2 | Elements, Compounds & Mixtures

1) Define a nucleus (1)

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2) Describe the relation of the mass number and the number of neutrons.

(3)

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3) Define an isotope. (1)

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4) Define an electron, in terms of its relative charge and mass (1)

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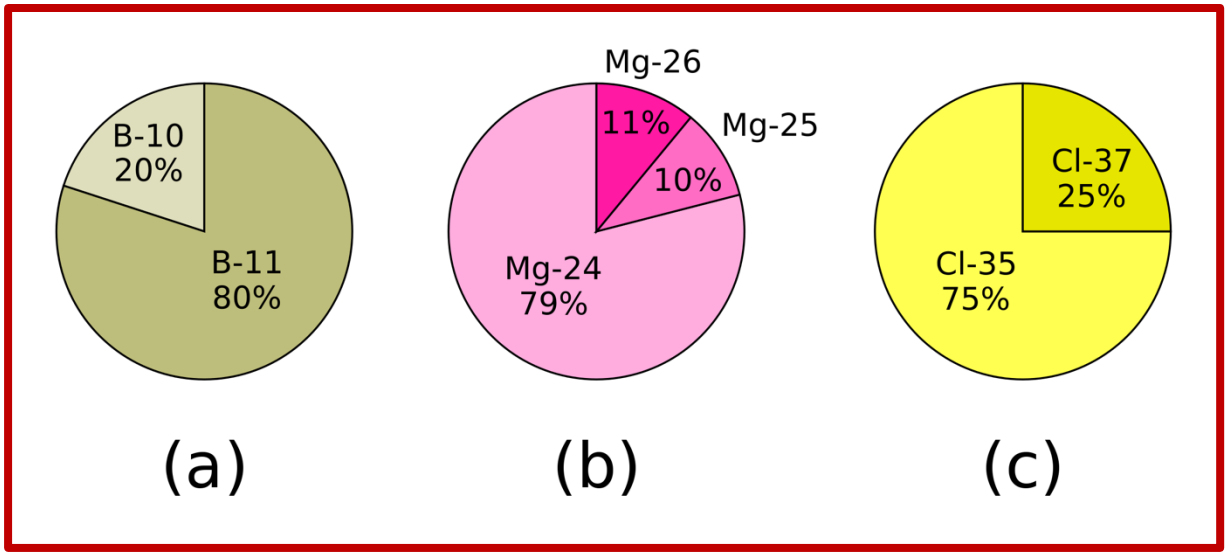
5) The following information is shown:



a) Explain the relation of an isotope's abundancy and its mass number.

(2)

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b) It is noted that Cl – 38 was discovered. If it takes 18% of the environment along with Cl-35 taking up 75%. What is the abundancy of Cl-37, now?

(3)

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6) Follow the following information and find the A_r:

Isotope A (i) – Mass number = 39
Abundancy = 89%

Isope A (ii) – Mass number = 40
Abundancy – 10%

Isotope A (iii) – Mass number = 41
Abundancy – 1%

(5)

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7) Which of the following best describes a mixture? (1)

A) Chemically combined elements
B) Not chemically combined
C) Homogenous
D) Heterogenous

8) Describe the process of crystallization (3)

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Total – 20 marks