

Samana TRUST

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Department of Science and Humanities Applied Chemistry Laboratory

Subject: Engineering Chemistry

Observation

Weight of empty crucible <u>22.4017</u> gm (W1)

Weight of crucible + Sample (Before heating) = <u>a3.2261</u> gm (W2)

Weight of sample before drying = $\frac{23 \cdot 2261 - 22 \cdot 4017}{gm(W_2 - W_1)}$

= <u>0.8244</u> gm (W3)

Weight of crucible + sample (after heating) = <u>a3.0740</u> gm (W4)

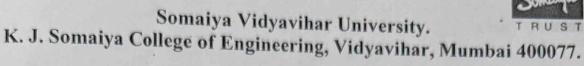
Weight of the sample (after heating) = $\frac{33.0740 - 33.4017}{\text{gm (W4-W1)}}$

$$= 0.6723 \text{ gm (W5)}$$

Loss in weight of sample = $\frac{0.6723 - 0.8244}{\text{gm (W5-W3)}}$

$$= 0.1521 \text{ gm (W6)}$$





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Calculations

Weight of sample = 0.8244 gm (W3)

Loss in weight = $0 \cdot 1521$ gm (W6)

%Moisture = Loss in weight
$$= 100$$
 Wt of sample taken $= 100$ W $= 100$ W

18. 4497

Result

: Percentage of moisture in given charcoal powder