

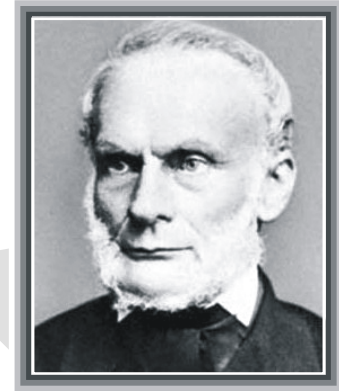
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NEET SYLLABUS

Thermal equilibrium and definition of temperature (zeroth law of Thermodynamics). Heat, temperature, thermal expansion; thermal expansion of solids, liquids, and gases. Anomalous expansion. Specific heat capacity: C_p , C_v - calorimetry; change of state - latent heat. Heat transfer- conduction and thermal conductivity, convection and radiation. Qualitative ideas of Black Body Radiation, Wein's displacement law Newton's law of cooling and Stefan's law. Kinetic theory of gases: Assumptions, concept of pressure. Kinetic energy and temperature; degrees of freedom, law of equipartition of energy (statement only) and application to specific heat capacities of gases; concept of mean free path. Heat, work and internal energy. First law of thermodynamics. Isothermal and adiabatic processes. Second law of the thermodynamics: Reversible and irreversible processes. Heat engines and refrigerators. Equation of state of a perfect gas, work done on compressing a gas.

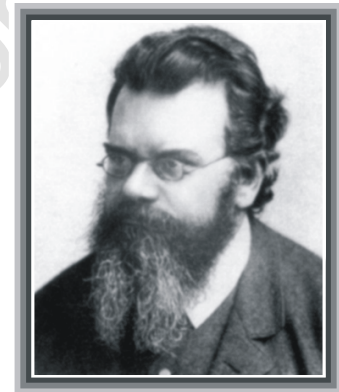
RUDOLF CLAUSIUS (1822-1888)

Rudolf Clausius, born in Poland, is generally regarded as the discoverer of the Second Law of Thermodynamics. Based on the work of Carnot and Thomson, Clausius arrived at the important notion of entropy that led him to a fundamental version of the Second Law of Thermodynamics that states that the entropy of an isolated system can never decrease. Clausius also worked on the kinetic theory of gases and obtained the first reliable estimates of molecular size, speed, mean free path, etc



LUDWIG BOLTZMANN (1844 – 1906)

Ludwig Boltzmann, born in Vienna, Austria, worked on the kinetic theory of gases independently of Maxwell. A firm advocate of atomism, that is basic to kinetic theory, Boltzmann provided a statistical interpretation of the Second Law of thermodynamics and the concept of entropy. He is regarded as one of the founders of classical statistical mechanics. The proportionality constant connecting energy and temperature in kinetic theory is known as Boltzmann's constant in his honour.



TEMPERATURE & THERMAL EXPANSION

1. TEMPERATURE & TEMPERATURE SCALES

• Temperature

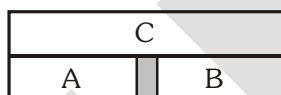
Temperature may be defined as the **degree of hotness or coldness** of a body. Heat energy flows from a body at higher temperature to that at lower temperature until their temperatures become equal. At this stage, the bodies are said to be in thermal equilibrium.

• Thermal equilibrium

Thermal equilibrium is a situation in which two objects would not exchange energy by heat or electromagnetic radiation if they were placed in thermal contact. Heat is the transfer of energy from one object to another object as a result of a difference in temperature between them.

• Zeroth law of thermodynamics

If objects A and B are separately in thermal equilibrium with a third object C (say thermometer), then objects A and B are in thermal equilibrium with each other. Zeroth law of thermodynamics introduces the concept of temperature. Two objects (or systems) are said to be in thermal equilibrium if their temperatures are same.



In measuring the temperature of a body, it is important that the thermometer should be in thermal equilibrium with the body whose temperature is to be measured.

• Measurement of Temperature

The branch of thermodynamics which deals with the measurement of temperature is called thermometry. A thermometer is a device used to measure the temperature of a body. The substances like liquids and gases which are used in the thermometer are called thermometric substances.

• Different Scales of Temperature

A thermometer can be graduated into following scales :

- The Centigrade or Celsius scale ($^{\circ}\text{C}$)
- The Fahrenheit scale ($^{\circ}\text{F}$)
- Kelvin scale (K)

• Comparison between Different Temperature Scales

The general formula for the conversion between different temperature scales is:

$$\frac{\text{K} - 273}{100} = \frac{\text{C}}{100} = \frac{\text{F} - 32}{180} = \frac{\text{X} - \text{LFP}}{\text{UFP} - \text{LFP}}$$

Where X \rightarrow Reading in unknown temperature scale, LFP \rightarrow Lower Fixed Point, UFP \rightarrow Upper Fixed Point

$$\text{Change in temperature} \quad \frac{\Delta \text{K}}{100} = \frac{\Delta \text{C}}{100} = \frac{\Delta \text{F}}{180} = \frac{\Delta \text{X}}{\text{UFP} - \text{LFP}}$$

GOLDEN KEY POINTS

- Although the temperature of a body can be raised without limit, it can not be lowered without limit and theoretically limiting low temperature is taken to be zero of the Kelvin scale (i.e. no negative temperature on Kelvin scale is possible).
- Though when universe was created 10^{10} years ago, its temp. was about 10^{39} K which at present is about 3K. The highest laboratory temperature is about 10^8 K (in fusion test reactor) while lowest 10^{-10} K (achieved in 1999 through nuclear spin cooling) Theory has established that zero Kelvin temperature can never be achieved practically.

Illustrations

Illustration 1.

Temperature of a patient is 40°C . Find the temperature on Fahrenheit scale ?

Solution :

$$\frac{F - 32}{180} = \frac{40 - 0}{100} \Rightarrow F = 104^\circ\text{F}$$

Illustration 2.

At what temperature is the Fahrenheit scale reading equal to twice of Celsius ?

Solution :

$$\frac{F - 32}{180} = \frac{C - 0}{100} \Rightarrow \frac{2x - 32}{180} = \frac{x - 0}{100} \Rightarrow x = 160$$

Illustration 3.

The lower and upper fixed points of a faulty thermometer are 5 W and 105 W. If the thermometer reads 25 W, what is the actual temperature in Celsius scale ?

Solution : $\frac{25 - 5}{100} = \frac{C - 0}{100} \Rightarrow C = 20^\circ\text{C}$

Illustration 4.

A thermometer with an arbitrary scale has the ice point at -20° and the steam point at 180° . When the thermometer reads 5° , a Centigrade thermometer will read

- (1) 7.5°C (2) 12.5°C (3) 16.5°C (4) -9.37°C

Solution :

$$\frac{C - 0}{100 - 0} = \frac{t - (-20)}{180 - (-20)} \quad (\text{Here } t = 5^\circ)$$

$$\Rightarrow \frac{C}{100} = \frac{5 + 20}{200} \Rightarrow C = 12.5^\circ\text{C}$$

Illustration 5.

The temperature of an iron piece is raised from 30°C to 90°C . What is the change in its temperature on the Fahrenheit scale and on the Kelvin scale?

Solution

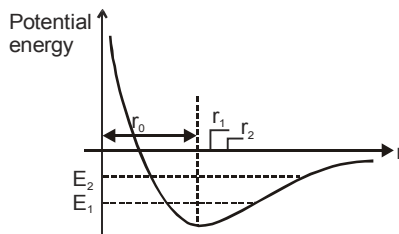
$$\Delta C = 90^\circ - 30^\circ = 60^\circ\text{C}$$

$$\text{Temperature difference on Fahrenheit Scale } \Delta F = \frac{9}{5} \Delta C = \frac{9}{5} (60^\circ\text{C}) = 108^\circ\text{F}$$

$$\text{Temperature difference on Kelvin Scale } \Delta K = \Delta C = 60\text{K}$$

2. THERMAL EXPANSION

When matter is heated without any change in its state, it usually expands. According to atomic theory of matter, asymmetry in potential energy curve is responsible for thermal expansion. As with rise in temperature the amplitude of vibration increases and hence energy of atoms increases, hence the average distance between the atom increases. So the matter as a whole expands.



Thermal expansion arises because the curve is not symmetrical about the equilibrium position r_0 . As the temperature rises the energy of the atom increases. The mean position when the energy is E_2 is not the same as that when the energy is E_1 .

- Thermal expansion is minimum in case of solids but maximum in case of gases because intermolecular forces are maximum in solids but minimum in gases.
- Solids can expand in one dimension (Linear expansion), two dimensions (Superficial expansion) and three dimensions (Volumetric expansion) while liquids and gases usually suffers change in volume only.

• Linear expansion :

$$\ell = \ell_0 (1 + \alpha \Delta\theta) \Rightarrow \Delta\ell = \ell_0 \alpha \Delta\theta$$

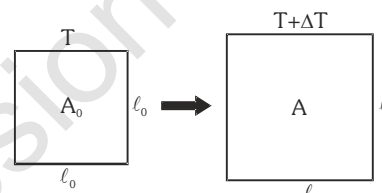


• Superficial (areal) expansion :

$$A = A_0 (1 + \beta \Delta\theta)$$

$$\text{Also } A_0 = \ell_0^2 \text{ and } A = \ell^2$$

$$\text{So } \ell^2 = \ell_0^2 (1 + \beta \Delta\theta) = [\ell_0 (1 + \alpha \Delta\theta)]^2 \Rightarrow \beta = 2\alpha$$



• Volumetric expansion :

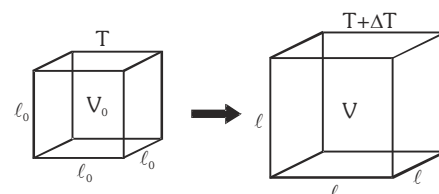
$$V = V_0 (1 + \gamma \Delta\theta) \text{ Also } V = \ell^3 \text{ and } V_0 = \ell_0^3 \text{ so } \gamma = 3\alpha$$

$$\ell^3 = [\ell_0 (1 + \alpha \Delta\theta)]^3 \Rightarrow 6\alpha = 3\beta = 2\gamma \text{ or } \alpha : \beta : \gamma = 1 : 2 : 3$$

α = coefficient of linear expansion

β = coefficient of superficial expansion

γ = coefficient of volumetric expansion

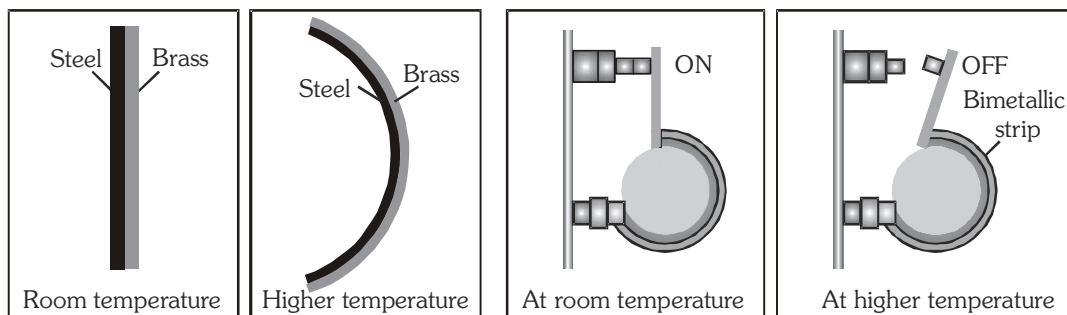


• Contraction on heating :

Some rubber like substances contract on heating because transverse vibration of atoms of substance dominate over longitudinal vibration which is responsible for expansion.

Applications of thermal Expansion in Solids

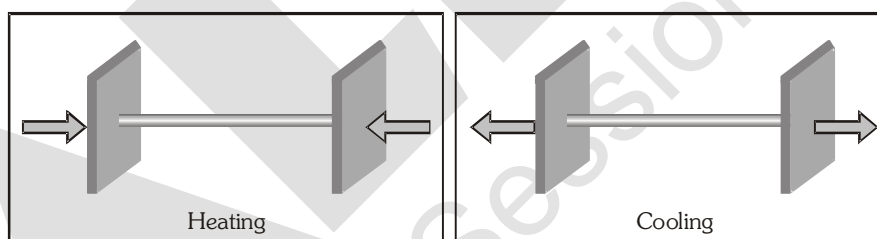
- Bi-metallic strip :** When two strips of equal length but of different materials (different coefficient of linear expansion) are joined together, it is called "Bi-metallic strip" and can be used in thermostat to break or make electrical contact. This strip has the characteristic property of bending on heating due to unequal linear expansion of the two metals. The strip will bend with metal of greater α on outer side. Coefficient of expansion is more for brass than steel.



- (b) **Effect of temperature on the time period of a simple pendulum :** A pendulum clock keeps proper time at temperature θ . If temperature is increased to θ' ($> \theta$) then due to linear expansion, length of pendulum increases and hence its time period will increase.

$$\text{Fractional change in time period } \frac{\Delta T}{T} = \frac{1}{2} \alpha \Delta \theta \quad (\because T \propto \sqrt{\ell} \quad \therefore \frac{\Delta T}{T} = \frac{1}{2} \frac{\Delta \ell}{\ell})$$

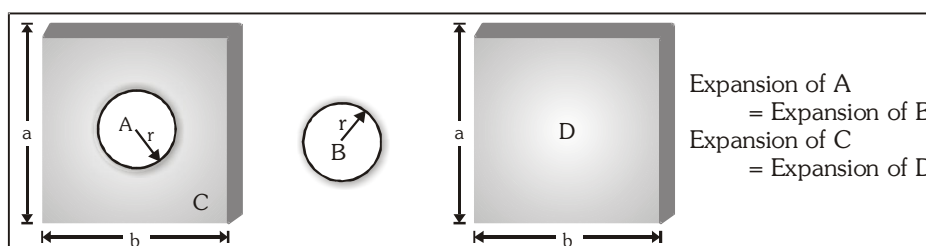
- Due to increment in its time period, a pendulum clock becomes slow in summer and will lose time. Loss of time in a time period $\Delta T = \frac{1}{2} \alpha \Delta \theta T$
 - The clock will lose time i.e. will become slow if $\theta' > \theta$ (in summer) and will gain time i.e. will become fast if $\theta' < \theta$ (in winter).
 - Since coefficient of linear expansion (α) is very small for invar, hence pendulums are made of invar to show the correct time in all seasons.
- (c) When a rod whose ends are rigidly fixed so as to prevent expansion or contraction undergoes a change in temperature, due to thermal expansion or contraction, a compressive or tensile stress is developed in it. Due to this thermal stress the rod will exert a large force on the supports. If the change in temperature of a rod of length L is $\Delta \theta$ then :-



$$\text{Thermal strain} = \frac{\Delta L}{L} = \alpha \Delta \theta \quad \therefore \alpha = \frac{\Delta L}{L} \times \frac{1}{\Delta \theta} \quad \text{So thermal stress} = Y \alpha \Delta \theta \quad \therefore Y = \frac{\text{stress}}{\text{strain}}$$

So force on the supports $F = Y A \alpha \Delta \theta$


- (d) **Expansion of cavity:** Thermal expansion of an isotropic object may be imagined as a photographic enlargement.

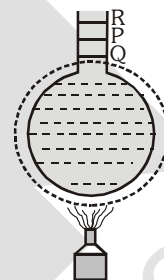


(e) Some other applications

- When rails are laid down on the ground, space is left between the ends of two rails
- The transmission cable are not tightly fixed to the poles
- Test tubes, beakers and cubicles are made up of pyrex-glass or silica because they have very low value of coefficient of linear expansion
- The iron rim to be put on a cart wheel is always of slightly smaller diameter than that of wheel
- A glass stopper jammed in the neck of a glass bottle can be taken out by warming the neck of the bottle.

- **Thermal Expansion in Liquids**

- Liquids do not have linear and superficial expansion but these only have volumetric expansion.
 - Since liquids are always to be heated along with a vessel which contains them so initially on heating the system (liquid + vessel), the level of liquid in vessel falls (as vessel expands more since it absorbs heat and liquid expands less) but later on, it starts rising due to faster expansion of the liquid.
- 



QP \rightarrow represents expansion of vessel

QR \rightarrow represents the real expansion of liquid.

- The actual increase in the volume of the liquid
= The apparent increase in the volume of liquid + the increase in the volume of the vessel.
- Liquids have two coefficients of volume expansion.

(i) Co-efficient of apparent expansion (γ_a)

It is due to apparent (that appears to be, but not in real) increase in the volume of liquid if expansion of vessel containing the liquid is not taken into account.

$$\gamma_a = \frac{\text{Apparent expansion in volume}}{\text{Initial volume} \times \Delta\theta} = \frac{(\Delta V)}{V \times \Delta\theta}$$

(ii) Co-efficient of real expansion (γ_r)

It is due to the actual increase in volume of liquid due to heating.

$$\gamma_r = \frac{\text{Real increase in volume}}{\text{Initial volume} \times \Delta\theta} = \frac{(\Delta V)}{V \times \Delta\theta}$$

- Also coefficient of expansion of flask $\gamma_{\text{vessel}} = \frac{(\Delta V)_{\text{vessel}}}{V \times \Delta \theta}$

$$\gamma_{\text{Real}} = \gamma_{\text{Apparent}} + \gamma_{\text{Vessel}}$$

- Change (apparent change) in volume in liquid relative to vessel is

$$\Delta V_{\text{app}} = V(\gamma_{\text{Real}} - \gamma_{\text{Vessel}}) \Delta\theta = V(\gamma_r - 3\alpha)\Delta\theta$$

α = Coefficient of linear expansion of the vessel.

- **Different level of liquid in vessel**

γ	ΔV	Level
$\gamma_{\text{Real}} > \gamma_{\text{Vessel}} (= 3\alpha) \Rightarrow \gamma_{\text{app}} > 0$	ΔV_{app} is positive	Level of liquid in vessel will rise on heating
$\gamma_{\text{Real}} < \gamma_{\text{Vessel}} (= 3\alpha) \Rightarrow \gamma_{\text{app}} < 0$	ΔV_{app} is negative	Level of liquid in vessel will fall on heating
$\gamma_{\text{Real}} = \gamma_{\text{Vessel}} (= 3\alpha) \Rightarrow \gamma_{\text{app}} = 0$	$\Delta V_{\text{app}} = 0$	Level of liquid in vessel will remain same

GOLDEN KEY POINTS

	K	C	F
Water boils	373.15	100	212
body temp.	310.2	37.0	98.6
Room temp.	300	27	80.6
Triple point of water	273.16	0.01	
Water freezes	273.15	0	32
Solid CO ₂	195	-78	-109
Hydrogen boils	20.7	-252.5	-422.5
Absolute zero	0	-273.15	-489.67

- Actually thermal expansion is always 3-D expansion. When other two dimensions of object are negligible with respect to one, then observations are significant only in one dimension and it is known as linear expansion.
- Relation between α , β and γ

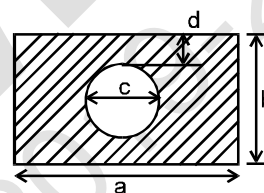
(i) For isotropic solids: $\alpha : \beta : \gamma = 1 : 2 : 3$ or $\frac{\alpha}{1} = \frac{\beta}{2} = \frac{\gamma}{3}$

(ii) For non-isotropic solids $\beta = \alpha_1 + \alpha_2$ and $\gamma = \alpha_1 + \alpha_2 + \alpha_3$. Here α_1 , α_2 and α_3 are coefficient of linear expansion in X, Y and Z direction.

Illustrations

Illustration 6.

A rectangular plate has a circular cavity as shown in the figure. If we increase its temperature then which dimension will increase in following figure.



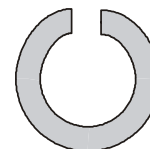
Solution :

Distance between any two point on an object increases with increase in temperature.

So, all dimension a, b, c and d will increase

Illustration 7.

A small ring having small gap is shown in figure on heating what will happen to the size of gap.



Solution :

Gap will also increase. The reason is same as in above example.

Illustration 8.

What is the percentage change in length of 1m iron rod if its temperature changes by 100°C.

α for iron is $2 \times 10^{-5}/^\circ\text{C}$.

Solution :

percentage change in length due to temperature change

$$\% \ell = \frac{\Delta \ell}{\ell} \times 100 = \alpha \Delta \theta \times 100 = 2 \times 10^{-5} \times 100 \times 100 = 0.2\%$$

Illustration 9.

A concrete slab has a length of 10 m on a winter night when the temperature is 0°C. Find the length of the slab on a summer day when the temperature is 35°C. The coefficient of linear expansion of concrete is $1.0 \times 10^{-5}/^\circ\text{C}$.

Solution :

$$\ell_t = 10(1 + 1 \times 10^{-5} \times 35) = 10.0035 \text{ m}$$

Illustration 10.

A glass vessel of volume 100 cm^3 is filled with mercury and is heated from 25°C to 75°C . What volume of mercury will overflow? Coefficient of linear expansion of glass = $1.8 \times 10^{-6}/^\circ\text{C}$ and coefficient of volume expansion of mercury is $1.8 \times 10^{-4}/^\circ\text{C}$.

Solution :

$$\Delta V = V_0(\gamma_L - \gamma_C) \Delta T = 100 \times \{1.8 \times 10^{-4} - 3 \times 1.8 \times 10^{-6}\} \times 50 \Rightarrow \Delta V = 0.87 \text{ cm}^3 \quad \text{Ans.}$$

Illustration 11.

There are two spheres of same radius and material at same temperature but one being solid while the other hollow. Which sphere will expand more if they are heated to the same temperature,

Solution

As thermal expansion of isotropic solids is similar to true photographic enlargement, expansion of a cavity is same as if it had been a solid body of the same material

i.e. $\Delta V = V\gamma \Delta\theta$

As here V , γ and $\Delta\theta$ are same for both solid and hollow spheres treated (cavity) ; so the expansion of both will be equal.



Illustration 12.

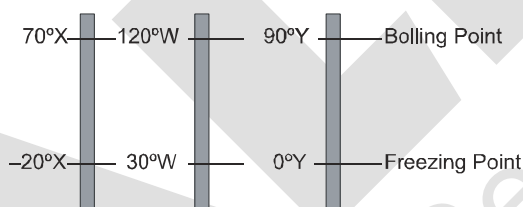
A steel wire of cross-sectional area 0.5 mm^2 is held between two fixed supports. If the wire is just taut at 20°C , determine the tension when the temperature falls to 0°C . Coefficient of linear expansion of steel is $1.2 \times 10^{-5}/^\circ\text{C}$ and its Young's modulus is $2.0 \times 10^{11} \text{ N/m}^2$.

Solution : Here final length is less than the original length so that strain is tensile and tensile force is given by

$$F = AY\alpha\Delta\theta = 0.5 \times 10^{-6} \times 2 \times 10^{11} \times 1.2 \times 10^{-5} \times 20 = 24 \text{ N}$$

BEGINNER'S BOX-1

- Write down the following temperatures in the increasing order 50°F , 50°C and 50 K .
- The figure shows three temperature scales with the freezing and boiling points of water indicated.



- Rank the size of a degree on these scales, greatest first.
 - Rank the following temperatures, highest first : 50°X , 50°W and 50°Y .
- What is the temperature at which we get the same reading on both the Centigrade and Fahrenheit scales?
 - A thin copper wire of length L increases in length by 1% when heated from temperature T_1 to T_2 . What is the percentage change in area when a thin copper plate having dimensions $2L \times L$ is heated from T_1 to T_2 ?
(A) 1% (B) 2% (C) 3% (D) 4%
 - A hole is drilled in a copper sheet. The diameter of the hole is 4.24 cm at 27.0°C . What is the change in the diameter of the hole when the sheet is heated to 227°C ?
Coefficient of linear expansion of copper = $1.70 \times 10^{-5} ^\circ\text{C}^{-1}$,
 - If a bimetallic strip is heated, it will
(A) bend towards the metal with lower thermal expansion coefficient.
(B) bend towards the metal with higher thermal expansion coefficient.
(C) twist itself into helix.
(D) have no bending.

HEAT

When a hot body is placed in contact with a cold one, the former gets colder and the latter warmer. From this observation it is natural to conclude that a certain quantity of heat has passed from the hot body to the cold one. Heat is a form of energy.

Heat is felt by its effects. Some of the effects of heat are :

- (a) Change in the degree of hotness
- (b) Change in length, surface area and volume
- (c) Change in state of a substance
- (d) Change in the resistance of a conductor
- (e) Thermo e.m.f. effect

SI Unit : J (joule). Also measured in calorie.

- **Calorie**

It is defined as the amount of heat required to raise the temperature of 1 g water by 1°C or 1 K.

- **International calorie**

International calorie is the amount of heat required to raise the temperature of 1g water from 14.5 °C to 15.5 °C rise of temperature at pressure of 1 atm.

- **kilo Calorie**

kilo calorie is defined as the amount of heat required to raise the temperature of 1 kg water from 14.5 °C to 15.5 °C. (1 kcal = 1000 calorie) at pressure of 1 atm.

- **British thermal unit (B.T.U.)**

It is the amount of heat required to raise the temperature of one pound water by 1°F. (1 B.T.U. = 252 calorie).

Mechanical equivalent of heat

According to Joule, work may be converted into heat and vice-versa. The ratio of work done to the heat produced is always constant. $\frac{W}{H} = \text{constant (J)} \Rightarrow W = J H$

W must be in joule, irrespective of nature of energy or work and H must be in calorie.

J is called mechanical equivalent of heat. It is not a physical quantity but simply a conversion factor.

It converts unit of work into that of heat and vice-versa.

$J = 4.186 \text{ joule/cal}$ or $4.186 \times 10^3 \text{ joule/kcal}$. For rough calculations we take $J = 4.2 \text{ joule/cal}$

Specific heat (s or c)

The amount of energy required to raise the temperature of unit mass of a substance by 1°C (or 1K) is called its specific heat. It is represented by s or c.

If the temperature of a substance of mass m changes from T to T + dT when it exchanges an amount of heat dQ with its surroundings then its specific heat is $c = \frac{1}{m} \frac{dQ}{dT}$

The specific heat depends on the pressure, volume and temperature of the substance.

For liquids and solids, specific heat measurements are most often made at a constant pressure as functions of temperature, because constant pressure is quite easy to produce experimentally.

Unit : joule/kg-K, cal/g -°C

Specific heat of water : $c_{\text{water}} = 1 \text{ cal/g-}^\circ\text{C} = 1 \text{ cal/g-K} = 1 \text{ kcal/kg-K} = 4200 \text{ joule/kg-K}$

When a substance does not undergo a change of state (i.e., liquid remains liquid or solid remains solid), then the amount of heat required to raise the temperature of mass m of the substance by an amount $\Delta\theta$ is

$$Q = ms\Delta\theta$$

Variation in specific heat of water is less than 1% over the interval from 0 to 100°C. Such a small variation is typical for most solids and liquids, so their specific heats can generally be taken to be constant over fairly large temperature ranges.

- There are many possible processes to give heat to a gas.

A specific heat can be associated to each such process which depends on the nature of process.

- Value of specific heats of gas can vary from zero (0) to infinity.
- Generally two types of specific heats are defined for a gas –
(a) Specific heat at constant volume (C_v) (b) Specific heat at constant pressure (C_p)
- These specific heats can be molar or gram.

Molar heat capacity

The amount of energy needed to raise the temperature of one mole of a substance by 1°C (or 1K) is called molar heat capacity. The molar heat capacity is the product of molecular weight and specific heat i.e.,

$$\text{Molar heat capacity } C = \text{Molecular weight } (M) \times \text{Specific heat } (c) \Rightarrow C = \frac{1}{\mu} \left(\frac{dQ}{dT} \right)$$

If the molecular mass of the substance is M and the mass of the substance is m then number of moles of the substance $\mu = \frac{m}{M} \Rightarrow C = \frac{M}{m} \left(\frac{dQ}{dT} \right)$

SI Unit : J/mol-K

Thermal capacity (Heat capacity)

The quantity of heat required to raise the temperature of the whole substance through 1°C is called thermal capacity. The thermal capacity of mass m of the whole substance of specific heat (s) = ms

Thermal capacity = mass \times specific heat

Thermal capacity depends on property of material of the body and mass of the body.

SI Unit : cal/°C or cal/K,

Dimensions : $ML^2 T^{-2}K^{-1}$

Water equivalent of a body

As the specific heat of water is unity so the thermal capacity of a body (ms) represents its water equivalent also.

- Mass of water having the same thermal capacity as the body is called the water equivalent of the body
- The water equivalent of a body is the amount of water that absorbs or gives out the same amount of heat as is done by the body when heated or cooled through 1°C.

Water equivalent = mass of body \times specific heat of the material $\Rightarrow (w = ms)$.

Latent heat or Hidden heat

When phase of a body changes, change of phase takes place at constant temperature [melting point or boiling point] and heat released or absorbed is $Q = mL$ where L is latent heat. Heat is absorbed if solid converts into liquid (at melting point) or liquid convert into vapours (at boiling point) and heat is released if liquid converts into solid or vapours convert into liquid.

- **Latent heat of fusion**

It is the quantity of heat (in kilocalories) required to change 1 kg mass of a substance from solid to liquid state at its melting point. Latent heat of fusion for ice : 80 kcal/kg = 80 cal /g.

- **Latent heat of vaporization**

The quantity of heat required to change its 1 kg mass from liquid to vapour state at its boiling point.

Latent heat of vaporisation for water : 536 kcal/kg = 536 cal/g = 540 cal/g

Change of State

- **Melting**

Conversion of solid into liquid state at constant temperature is known as melting.

- **Boiling**

Evaporation within the whole mass of the liquid is called boiling. Boiling takes place at a constant temperature known as boiling point. A liquid boils when the saturated vapour pressure on its surface is equal to atmospheric pressure. Boiling point reduces on decreasing pressure.

- **Evaporation**

Conversion of liquid into vapours at all temperatures is called evaporation. It is a surface phenomenon. Greater the temperature, faster is the evaporation. Smaller the boiling point of liquid, more rapid is the evaporation. Smaller the humidity, more is the evaporation. Evaporation increases on decreasing pressure that is why evaporation is faster in vacuum.

- **Heat of evaporation**

Heat required to change unit mass of a liquid into vapour at a given temperature is called heat of evaporation at that temperature.

- **Sublimation**

Direct conversion of solid into vapour state is called sublimation.

- **Heat of sublimation**

Heat required to change unit mass of solid directly into vapours at a given temperature is called heat of sublimation at that temperature.

- Camphor and ammonium chloride sublimes on heating in normal conditions.

- A block of ice sublimes into vapours on the surface of moon because of very-very low pressure on its surface

- **Condensation**

The process of conversion from gaseous or vapour state to liquid state is known as condensation .

These materials again get converted to vapour or gaseous state on heating.

- **Hoar frost**

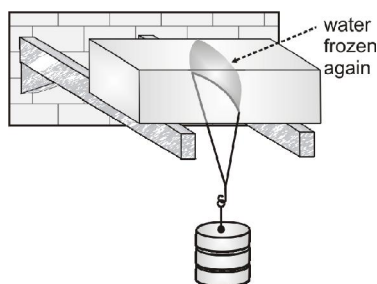
Direct conversion of vapours into solid is called hoar frost. This process is just reverse of the process of sublimation.

Ex. : Formation of snow by freezing of clouds.

- **Regelation**

Regelation is the melting of ice caused by pressure and its resolidification when the pressure is removed. Ice shrinks when it melts, and if pressure is applied, deliberately promoting shrinkage, it is found that melting is thereby assisted. In other words, melting of cold ice is ordinarily effected by raising the temperature, but if pressure is present to help with the shrinkage the temperature need not be raised so much.

Ice heals up after being cut through by the wire. Melting takes place under the wire because pressure lowers the melting temperature. Refreezing (regelation) occurs above the wire when the water escapes to normal pressure again.



Increase of pressure lowers the melting (or freezing) point of water. Conversely, if a substance expands on melting, the melting point is raised by pressure.

Phase of a substance

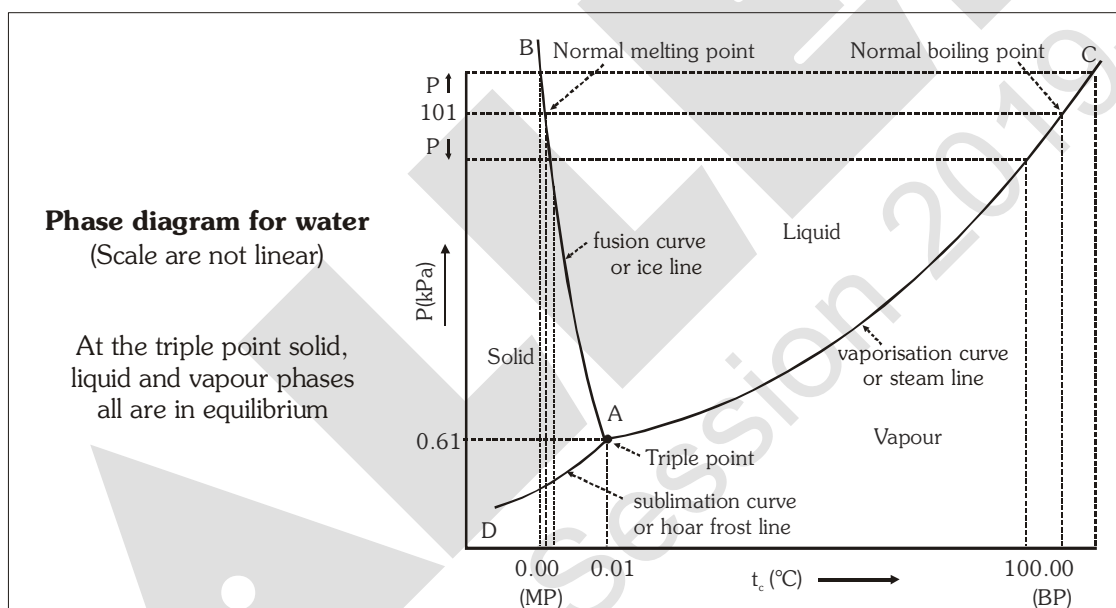
The phase of a substance is defined as its form which is homogeneous, physically distinct and mechanically separable from the other forms of that substance.

Phase diagram

- A phase diagram is a graph in which pressure (P) is represented along the y-axis and temperature (T) is represented along the x-axis.
- **Characteristics of Phase diagram**
 - Different phases of a substance can be shown on a phase diagram.
 - A region on the phase diagram represents a single phase of the substance, a curve represents equilibrium between two phases and a common point represents equilibrium between three phases.
 - A phase diagram helps to determine the condition under which the different phases are in equilibrium.
 - A phase diagram is useful for finding a convenient way in which a desired change of phase can be produced.

Phase diagram for water

The phase diagram for water consists of three curves AB, AC and AD meeting each other at the point A, these curves divide the phase diagram into three regions.



Region to the left of the curve AB and above the curve AD represents the solid phase of water (ice). The region to the right of the curve AB and above the curve AC represents the liquid phase of water. The region below the curves AC and AD represents the gaseous phase of water (i.e. water vapour). A curve on the phase diagram represents the boundary between two phases of the substance.

Along any curve the two phases can coexist in equilibrium

- Along curve AB, ice and water can remain in equilibrium. This curve is called fusion curve or ice line. This curve shows that the melting point of ice decreases with increase in pressure.
- Along the curve AC, water and water vapour can remain in equilibrium. The curve is called vaporisation curve or steam line. The curve shows that the boiling point of water increases with increase in pressure.
- Along the curve AD, ice and water vapour can remain in equilibrium. This curve is called sublimation curve or hoar frost line.

Triple point of water

The three curves in the phase diagram of water meet at a single point A, which is called the triple point of water. The triple point of water represents the co-existence of all the three phases of water ice water and water vapour in equilibrium. The pressure corresponding to triple point of water is 6.03×10^{-3} atmosphere or 4.58 mm of Hg and temperature corresponding to it is 273.16K.

- Significance of triple point of water**

Triple point of water represents a unique condition and it is used to define the absolute temperature. While making Kelvin's absolute scale, upper fixed point is 273.16 K and lower fixed point is 0 K. One kelvin of temperature is $\frac{1}{273.16}$ part of the temperature of triple point of water.

Effect of change in pressure on M.P. and B.P. for water

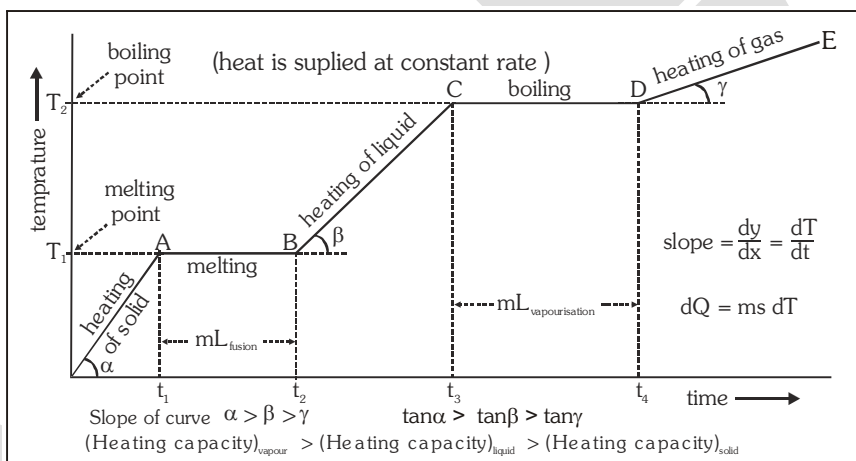
- If $P \uparrow$ $\xrightarrow{\text{then result}}$ M.P. \downarrow & B.P. \uparrow
- If $P \downarrow$ $\xleftarrow{\text{then result}}$ M.P. \uparrow & B.P. \downarrow

Example :

- (1) A bottle is filled with water at 30°C on opening at moon then water will boil and vapourised
- (2) At higher altitudes of mountain, food can not be cooked properly.

Heating curve

If to a given mass (m) of a solid, heat is supplied at constant rate and a graph is plotted between temperature and time, as shown in figure, it is called heating curve.



- In the region OA**

Rate of heat supply P is constant and temperature of solid is changing with time.

$$\text{So, } Q = mc_s \Delta T \Rightarrow P \Delta t = mc_s \Delta T \quad [\because Q = P \Delta t]$$

$\therefore \frac{\Delta T}{\Delta t} =$ The slope of temperature-time curve so specific heat of solid $c_s \propto \frac{1}{\text{slope of line OA}}$ specific heat (or thermal capacity) is inversely proportional to the slope of temperature-time curve.

- In the region AB**

Temperature is constant, so it represents change of state, i.e., melting of solid at melting point T_1 . At point A melting starts and at point B all solid is converted into liquid. So between A and B substance is partly solid and partly liquid. If L_F is the latent heat of fusion then

$$Q = mL_F \Rightarrow L_F = \frac{P(t_2 - t_1)}{m} \quad [\text{as } Q = P(t_2 - t_1)] \Rightarrow L_F \propto \text{length of line AB}$$

i.e., Latent heat of fusion is proportional to the length of line of zero slope.

$$[\text{In this region specific heat} \propto \frac{1}{\tan 0^\circ} = \infty]$$

• **In the region BC**

Temperature of liquid increases so specific heat (or thermal capacity) of liquid will be inversely proportional

to the slope of line BC, $c_L \propto \frac{1}{\text{slope of line BC}}$

• **In the region CD**

Temperature is constant, so it represents change of state, i.e., liquid is boiling at boiling point T_2 . At C all substance is in liquid state while at D is vapour state and between C and D partly liquid and partly gas. The length of line CD is proportional to latent heat of vaporisation, i.e., $L_v \propto \text{Length of line CD}$.

[In this region specific heat $\propto \frac{1}{\tan 0^\circ} = \infty$]

• **In the region DE**

The line DE represents gaseous state of substance with its temperature increasing linearly with time. The reciprocal of slope of line will be proportional to specific heat or thermal capacity of substance in vapour state.

Law of Mixtures : Principle of calorimetry

- When two bodies at different temperatures are mixed, heat will be transferred from body at higher temperature to a body at lower temperature till both acquire same temperature. The body at higher temperature releases heat while body at lower temperature absorbs it, so that

$$\boxed{\text{Heat lost} = \text{Heat gained}} \quad (\text{Principle of calorimetry})$$

Principle of calorimetry represents the law of conservation of heat energy.

- Temperature of mixture (T) is always \geq lower temperature (T_L) and \leq higher temperature (T_H), $\boxed{T_L \leq T \leq T_H}$

The temperature of mixture can never be lesser than lower temperature (as a body cannot be cooled below the temperature of cooling body) and greater than higher temperature (as a body cannot be heated above the temperature of heating body). Further more usually, rise in temperature of one body may not be equal to the fall in temperature of the other body though heat gained by one body is equal to the heat lost by the other.

Illustrations

Illustration 13.

5 g ice at 0°C is mixed with 5 g of steam at 100°C . What is the final temperature?

Solution

Heat required by ice to raise its temperature to 100°C ,

$$Q_1 = m_1 L_1 + m_1 c_1 \Delta\theta_1 = 5 \times 80 + 5 \times 1 \times 100 = 400 + 500 = 900 \text{ cal}$$

$$\text{Heat given by steam when condensed } Q_2 = m_2 L_2 = 5 \times 536 = 2680 \text{ cal}$$

As $Q_2 > Q_1$. This means that whole steam is not even condensed.

Hence temperature of mixture will remain at 100°C .

Illustration 14.

A calorimeter of heat capacity 100 J/K is at room temperature of 30°C . 100 g of water at 40°C of specific heat 4200 J/kg-K is poured into the calorimeter. What is the temperature of water in calorimeter?

Solution

Let the temperature of water in calorimeter is t . Then heat lost by water = heat gained by calorimeter

$$(0.1) \times 4200 \times (40 - t) = 100 (t - 30) \Rightarrow 420 \times 40 - 420 t = 100 t - 3000 \Rightarrow t = 38.07^\circ\text{C}$$

Illustration 15.

Find the quantity of heat required to convert 40 g of ice at -20°C into water at 20°C .

Given $L_f = 0.336 \times 10^6 \text{ J/kg}$. Specific heat of ice = 2100 J/kg-K , specific heat of water = 4200 J/kg-K

Solution

Heat required to raise the temperature of ice from -20°C to 0°C = $0.04 \times 2100 \times 20 = 1680 \text{ J}$

Heat required to convert the ice into water at 0°C = $mL = 0.04 \times 0.336 \times 10^6 = 13440 \text{ J}$

Heat required to heat water from 0°C to 20°C = $0.04 \times 4200 \times 20 = 3360 \text{ J}$

Total heat required = $1680 + 13440 + 3360 = 18480 \text{ J}$

Illustration 16.

Steam at 100°C is passed into 1.1 kg of water contained in a calorimeter of water equivalent 0.02 kg at 15°C till the temperature of the calorimeter and its contents rises to 80°C . What is the mass of steam condensed?

Latent heat of steam = 536 cal/g .

Solution

Heat required by (calorimeter + water)

$$Q = (m_1 c_1 + m_2 c_2) \Delta\theta = (0.02 + 1.1 \times 1) (80 - 15) = 72.8 \text{ kcal}$$

If m is mass of steam condensed, then heat given by steam

$$Q = mL + mc \Delta\theta = m \times 536 + m \times 1 \times (100 - 80) = 556 m \quad \therefore 556 m = 72.8$$

$$\therefore \text{Mass of steam condensed } m = \frac{72.8}{556} = 0.130 \text{ kg}$$

Illustration 17.

An iron block of mass 2 kg, fall from a height 10 m. After colliding with the ground it loses 25% energy to surroundings. Then find the temperature rise of the block. (Take specific heat of iron $470 \text{ J/kg } ^{\circ}\text{C}$)

Solution :

$$mS\Delta\theta = \frac{3}{4} mgh \quad \Rightarrow \quad \Delta\theta = \frac{3 \times 10 \times 10}{4 \times 470} = 0.159^{\circ}\text{C}$$

Illustration 18.

The temperature of equal masses of three different liquids A, B, and C are 10°C , 15°C and 20°C respectively. The temperature when A and B are mixed is 13°C and when B and C are mixed, it is 16°C . What will be the temperature when A and C are mixed?

Solution :

when A and B are mixed

$$mS_1 \times (13 - 10) = m \times S_2 \times (15 - 13)$$

$$3S_1 = 2S_2 \quad \dots\dots(1)$$

when B and C are mixed

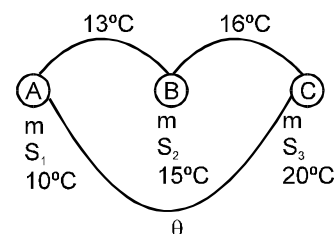
$$S_2 \times 1 = S_3 \times 4 \quad \dots\dots(2)$$

when C and A are mixed

$$S_1(\theta - 10) = S_3 \times (20 - \theta) \quad \dots\dots(3)$$

by using equation (1), (2) and (3)

$$\text{we get } \theta = \frac{140}{11}^{\circ}\text{C}$$



MODES OF HEAT TRANSFER

Heat is a form of energy which transfers from a body at higher temperature to a body at lower temperature.

The transfer of heat from one body to another may take place by any one of the following modes :

- **Conduction**

The process in which the material takes an active part by molecular action and energy is passed from one particle to another is called conduction. It is predominant in solids.

- **Convection**

The transfer of energy by actual motion of particle of medium from one place to another is called convection. It is predominant in fluids (liquids and gases).

- **Radiation**

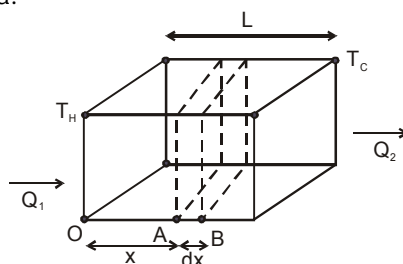
Quickest way of transmission of heat is known as radiation. In this mode of energy transmission, heat is transferred from one place to another without effecting the intervening medium.

Conduction	Convection	Radiation
Heat Transfer due to Temperature difference	Heat transfer due to density difference	Heat transfer without any medium
Due to free electron or vibration motion of molecules	Actual motion of particles	Electromagnetic radiation
Heat transfer in solids and Hg	Heat transfer in fluids (Liquid + gas)	No medium required
Slow process	Slow process	Fast process (3×10^8 m/sec)
Irregular path	Irregular path	Straight line (like light)

3.1 Thermal conduction

The process by which heat is transferred from hot part to cold part of a body through the transfer of energy from one particle to another particle of the body without the actual movement of the particles from their equilibrium positions is called conduction. The process of conduction takes place only in solid body (except Hg). Heat transfer by conduction from one part of body to another continues till their temperatures become equal.

For example if you hold an iron rod with one of its end on a fire for some time, the handle will become hot. The heat is transferred from the fire to the handle by conduction along the length of iron rod. The vibrational amplitude of atoms and electrons of the iron rod at the hot end takes relatively higher values due to the higher temperature of their environment. These increased vibrational amplitude are transferred along the rod, from atom to atom during collision between adjacent atoms. In this way a region of rising temperature extends itself along the rod to your hand.



Consider a slab of face area A , Lateral thickness L , whose faces have temperatures T_H and T_C ($T_H > T_C$). Now consider two cross sections in the slab at positions A and B separated by a lateral distance of dx .

Let temperature of face A be T and that of face B be $T + \Delta T$. Then experiments show that Q , the amount of heat crossing the area A of the slab at position x in time t is given by

$$\boxed{\frac{Q}{t} = -KA \frac{dT}{dx}} \quad \text{or} \quad \boxed{\frac{Q}{t} = \frac{KA(T_H - T_C)}{L}}$$

$K \rightarrow$ Thermal conductivity

\rightarrow It is the measure of the ability of material to conduct the heat.

Here K is a constant depending on nature of the material of the slab and is named thermal conductivity of

the material, and the quantity $\left(\frac{dT}{dx}\right)$ is called temperature gradient. The $(-)$ sign shows heat flows from high temperature to low temperature (ΔT is a $-ve$ quantity).

Steady State

If the temperature of a cross-section at any position x in the above slab remains constant with time (remember, it does vary with position x), the slab is said to be in steady state and temperature of rod is not same.

Remember steady-state is distinct from thermal equilibrium for which temperature at any position (x) in the slab must be same.

For a conductor in steady state there is no absorption or emission of heat at any cross-section. (as temperature at each point remains constant with time). The left and right faces are maintained at constant temperatures T_H and T_C respectively, and all other faces must be covered with adiabatic walls so that no heat escapes through them and same amount of heat flows through each cross-section in a given interval of time.

Hence $Q_1 = Q = Q_2$. Consequently the temperature gradient is constant throughout the slab.

Hence,
$$\frac{dT}{dx} = \frac{\Delta T}{L} = \frac{T_f - T_i}{L} = \frac{T_C - T_H}{L}$$

and
$$\frac{Q}{t} = -KA \frac{\Delta T}{L} \Rightarrow \frac{Q}{t} = KA \left(\frac{T_H - T_C}{L} \right)$$

Here Q is the amount of heat flowing through a cross-section of slab at any position in a time interval t .

Thermal conductivity (K) :

- It depends on nature of material.

Order of thermal conductivity $\boxed{\text{Ag} > \text{Cu} > \text{Au} > \text{Al}}$

$K \begin{cases} \text{For Ag maximum is } (410 \text{ W/mK}) \\ \text{For Freon minimum is } 12 (0.008 \text{ W/mK}) \end{cases}$

- SI Unit** : $\text{J s}^{-1} \text{ m}^{-1} \text{ K}^{-1}$ **Dimensions** : $\text{M}^1 \text{ L}^1 \text{ T}^{-3} \theta^{-1}$
- For an ideal or perfect conductor of heat the value of $K = \infty$
- For an ideal or perfect bad conductor or insulator the value of $K = 0$
- For cooking the food, low specific heat and high conductivity utensils are most suitable.

Application of Thermal Conduction

- In winter, the iron chairs appear to be colder than the wooden chairs.
- Cooking utensils are made of aluminium and brass whereas their handles are made of wood.
- Ice is covered in gunny bags to prevent melting of ice.
- We feel warm in woollen clothes and fur coat.
- Two thin blankets are warmer than a single blanket of double the thickness.
- Birds often swell their feathers in winter.
- A new quilt is warmer than old one.

Thermal Resistance to conduction

If you are interested in insulating your house from cold weather or for that matter keeping the meal hot in your tiffin-box, you are more interested in poor heat conductors, rather than good conductors. For this reason, the concept of thermal resistance R has been introduced.

For a slab of cross-section A , Lateral thickness L and thermal conductivity K ,
$$R = \frac{L}{KA}$$

In terms of R , the amount of heat flowing through a slab in steady-state (in time t)
$$\frac{Q}{t} = \frac{(T_H - T_L)}{R}$$

If we name $\frac{Q}{t}$ as thermal current i_T then,

$$i_T = \frac{T_H - T_L}{R}$$

This is mathematically equivalent to OHM's law, with temperature doing the role of electric potential. Hence results derived from OHM's law are also valid for thermal conduction.

More over, for a slab in steady state we have seen earlier that the thermal current i_T remains same at each cross-section. This is analogous to Kirchhoff's current law in electricity, which can be very conveniently applied to thermal conduction.

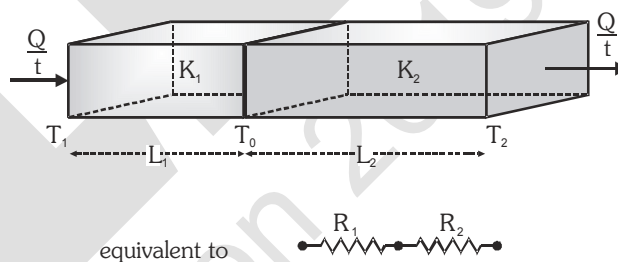
- Equivalent conductivity for Heat flow through slabs in series**

$$R_{eq} = R_1 + R_2$$

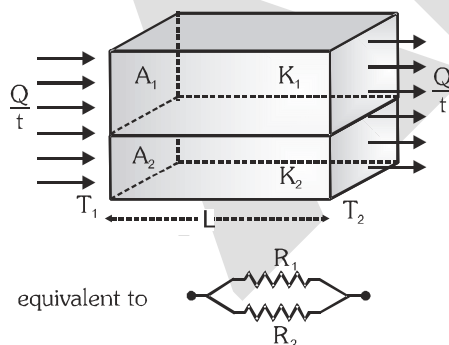
$$\frac{L_1 + L_2}{K_{eq}A} = \frac{L_1}{K_1A} + \frac{L_2}{K_2A}$$

Equivalent thermal conductivity of the system is

$$K_{eq} = \frac{L_1 + L_2}{\frac{L_1}{K_1} + \frac{L_2}{K_2}} = \frac{\sum L_i}{\sum \frac{L_i}{K_i}}$$



- Equivalent thermal conductivity for Heat flow through slabs in parallel**



$$R = \frac{L}{KA}$$

$$\frac{1}{R_{eq}} = \frac{1}{R_1} + \frac{1}{R_2}$$

$$\frac{K_{eq}}{L}(A_1 + A_2) = \frac{K_1 A_1}{L} + \frac{K_2 A_2}{L}$$

Equivalent thermal conductivity

$$K_{eq} = \frac{K_1 A_1 + K_2 A_2}{A_1 + A_2} = \frac{\sum K_i A_i}{\sum A_i}$$

Growth of Ice on Lakes

In winter atmospheric temperature falls below 0°C and water in the lake start freezing. Let at time t thickness of ice on the surface of the lake = x and air temperature = $-\theta^{\circ}\text{C}$

The temperature of water in contact with the lower surface of ice = 0°C

Let area of the lake = A

Heat escaping through ice in time dt is $dQ = KA \frac{[0 - (-\theta)]}{x} dt$

Due to escape of this heat increasing extra thickness of ice = dx

Mass of this extra thickness of ice is $m = \rho V = \rho A \cdot dx$

$$dQ = mL = (\rho A \cdot dx) L$$

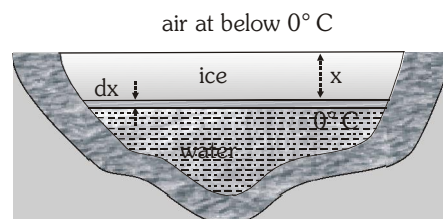
$$\therefore KA \frac{\theta}{x} dt = (\rho A \cdot dx) L \Rightarrow dt = \frac{\rho L}{K\theta} x dx$$

So time taken by ice to grow a thickness x is $t = \frac{\rho L}{K\theta} \int_0^x x dx = \frac{1}{2} \frac{\rho L}{K\theta} x^2$

So time taken by ice to grow from thickness x_1 to thickness x_2 is

$$t = t_2 - t_1 = \frac{1}{2} \frac{\rho L}{K\theta} (x_2^2 - x_1^2) \quad \text{and} \quad t \propto (x_2^2 - x_1^2)$$

Time taken to double and triple the thickness ratio $t_1 : t_2 : t_3 :: 1^2 : 2^2 : 3^2$ So $t_1 : t_2 : t_3 :: 1 : 4 : 9$



3.2 Convection

Convection requires a medium and is the process in which heat is transferred from one place to other by actual movement of heated substance (usually fluid). The type of convection which results from difference in densities is called natural convection (for example, a fluid in a container heated through its bottom). However, if a heated fluid is forced to move by a blower, fan or pump, the process is called forced convection.

Phenomena Based on convection :

(i) Land and sea breezes :

The heat from the Sun is absorbed more rapidly by land than by sea-water. Moreover, the specific heat of land is low as compared to that of sea-water. Consequently, the rise in temperature of land is higher as compared to that of sea-water. To sum-up, land is hotter than the sea during day time. As a result of this, the colder air over the sea blows towards the land. This is called sea-breeze.

At night, air blows from land towards sea. This is called land breeze.

(ii) Formation of trade winds :

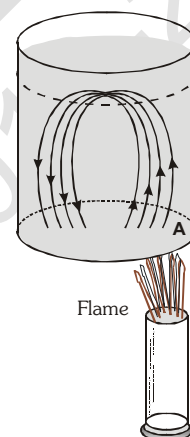
The surface of Earth near the equator gets heated strongly. So, the air in contact with the surface of Earth at the equator expands and rises upwards. As a result of this, a low pressure is created at the equator.

At the poles, the air in the upper atmosphere gets cooled and comes down. So, a high pressure is created at the poles. Due to difference of pressures at the poles and equator, the air at the poles moves towards the equator, rises up, moves towards poles and so on. In this way, a wind is formed in the atmosphere.

The rotation of the Earth also affects the motion of the wind. Due to anti-clockwise rotation of Earth the warm wind blowing from equator to north drifts towards east. The steady wind blowing from north-East to equator, near the surface of Earth, is called trade wind.

(iii) Monsoons :

In summer, the peninsular mass of central Asia becomes more strongly heated than the water of the Indian Ocean. This is due to the fact that the specific heat of water is much higher than that of the soil and rocks. Hot air from the heated land mass rises up and moves towards the Indian ocean. Air filled with moisture flows over the Indian ocean on the south towards heated land mass. When obstructed by mountains, the moist air rushes upwards to great height. In the process, it gets cooled. Consequently, the moisture condenses and falls as rain.



(iv) Ventillation :

Ventillators of exhaust fan in a room help to remove impure and warm air from a room. The fresh air from outside blows into the room. This is all due to the forced convection current set up in the room.

(v) To regulate temperature in the human body :

Heat transfer in the human body involves a combination of mechanisms. These together maintain a remarkably uniform temperature in the human body in spite of large changes in environmental conditions. The chief internal mechanism is forced convection. The heart serves as the pump and the blood as the circulating fluid.

Some important points :

- Natural convection takes place from bottom to top due to gravity while forced convection in any direction.
- In case of natural convection, convection currents move warm air upwards and cool air downwards. This is why heating is done from base, while cooling from the top.
- Natural convection is not possible in a gravity free region such as a freely falling lift or an orbiting satellite.
- Natural convection plays an important role in ventilation, in changing climate and weather and in forming land and sea breezes and trade winds.
- The forced convection of blood in our body by a pump (heart) helps in keeping the temperature of body constant.
- Convection takes place in fluids (gases and liquid)

GOLDEN KEY POINTS

- For heat propagation via natural convection, temperature gradient exists in vertical direction and not in horizontal direction.
- Most of heat transfer that is taking place on Earth is by convection, the contribution due to conduction and radiation is very small.

Illustrations**Illustration 20.**

One face of an aluminium cube of edge 2 metre is maintained at 100 °C and the other end is maintained at 0 °C. All other surfaces are covered by adiabatic walls. Find the amount of heat flowing through the cube in 5 seconds. (thermal conductivity of aluminium is 209 W/m-°C)

Solution

Heat will flow from the end at 100°C to the end at 0°C.

Area of cross-section perpendicular to direction of heat flow, $A = 4\text{m}^2$ then

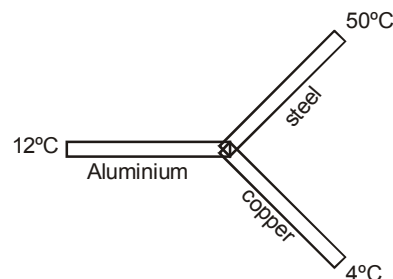
$$\frac{Q}{t} = KA \frac{(T_H - T_C)}{L}$$

$$Q = \frac{(209\text{ W/m}^\circ\text{C})(4\text{m}^2)(100^\circ\text{C} - 0^\circ\text{C})(5\text{ sec})}{2\text{m}} = 209\text{ kJ}$$

Illustration 21.

Three identical rods of length 1m each, having cross-section area of 1cm^2 each and made of Aluminium, copper and steel respectively are maintained at temperatures of 12°C, 4°C and 50°C respectively at their separate ends. Find the temperature of their common junction.

[$K_{\text{Cu}} = 400\text{ W/m-K}$, $K_{\text{Al}} = 200\text{ W/m-K}$, $K_{\text{steel}} = 50\text{ W/m-K}$]

**Solution**

$$R_{\text{Al}} = \frac{L}{KA} = \frac{1}{200 \times 10^{-4}} = \frac{10^4}{200}$$

$$\text{Similarly } R_{\text{steel}} = \frac{10^4}{50} \text{ and } R_{\text{copper}} = \frac{10^4}{400}$$

Let temperature of common junction = T

then from Kirchhoff's current law, $i_{Al} + i_{steel} + i_{Cu} = 0$

$$\Rightarrow \frac{T-12}{R_{Al}} + \frac{T-50}{R_{steel}} + \frac{T-4}{R_{Cu}} = 0$$

$$\Rightarrow (T-12) 200 + (T-50) 50 + (T-4) 400 = 0$$

$$\Rightarrow 4(T-12) + (T-50) + 8(T-4) = 0$$

$$\Rightarrow 13T = 48 + 50 + 32 = 130 \Rightarrow T = 10^\circ\text{C}$$

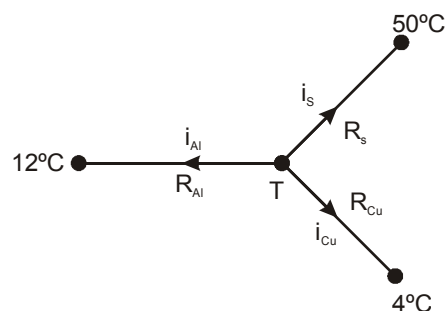


Illustration 22.

The thermal conductivity of brick is $1.7 \text{ W m}^{-1} \text{ K}^{-1}$, and that of cement is $2.9 \text{ W m}^{-1} \text{ K}^{-1}$. What thickness of cement will have same insulation as the brick of thickness 20 cm? Assuming their area to be same.

Solution

Since $Q = \frac{KA(T_1 - T_2)t}{L}$. For same insulation by the brick and cement; Q , A ($T_1 - T_2$) and t do not change.

Hence, $\frac{K}{L}$ remain constant. If K_1 and K_2 be the thermal conductivities of brick and cement respectively and

$$L_1 \text{ and } L_2 \text{ be the required thickness then } \frac{K_1}{L_1} = \frac{K_2}{L_2} \text{ or } \frac{1.7}{20} = \frac{2.9}{L_2} \quad \therefore L_2 = \frac{2.9}{1.7} \times 20 = 34.12 \text{ cm}$$

Illustration 23.

Two vessels of different materials are identical in size and wall-thickness. They are filled with equal quantities of ice at 0°C . If the ice melts completely, in 10 and 25 minutes respectively then compare the coefficients of thermal conductivity of the materials of the vessels.

Solution

Let K_1 and K_2 be the coefficients of thermal conductivity of the materials, and t_1 and t_2 be the time in which ice melts in the two vessels. Since both the vessels are identical, so A and L in both the cases is same.

$$\text{Now, } Q = \frac{K_1 A (\theta_1 - \theta_2) t_1}{L} = \frac{K_2 A (\theta_1 - \theta_2) t_2}{L} \Rightarrow \frac{K_1}{K_2} = \frac{t_2}{t_1} = \frac{25 \text{ min}}{10 \text{ min}} = \frac{5}{2}$$

Illustration 24.

Two plates of equal areas are placed in contact with each other. Their thickness are 2.0 cm and 5.0 cm respectively. The temperature of the external surface of the first plate is -20°C and that of the external surface of the second plate is 20°C . What will be the temperature of the contact surface if the plate (i) are of the same material, (ii) have thermal conductivities in the ratio 2 : 5.

Solution

$$\text{Rate of flow of heat in the plates is } \frac{Q}{t} = \frac{K_1 A (\theta_1 - \theta)}{L_1} = \frac{K_2 A (\theta - \theta_2)}{L_2} \dots (i)$$

(i) Here $\theta_1 = -20^\circ\text{C}$, $\theta_2 = 20^\circ\text{C}$,

$$L_1 = 2 \text{ cm} = 0.02 \text{ m}, L_2 = 5 \text{ cm} = 0.05 \text{ m and } K_1 = K_2 = K$$

$$\therefore \text{equation (i) becomes } \frac{KA(-20 - \theta)}{0.02} = \frac{KA(\theta - 20)}{0.05}$$

$$\therefore 5(-20 - \theta) = 2(\theta - 20) \Rightarrow -100 - 5\theta = 2\theta - 40 \Rightarrow 7\theta = -60 \Rightarrow \theta = -8.6^\circ\text{C}$$

$$(ii) \quad \frac{K_1}{K_2} = \frac{2}{5} \text{ or } K_1 = \frac{2}{5} K_2$$

$$\therefore \text{from equation (i) } \frac{2/5 K_2 A (-20 - \theta)}{0.02} = \frac{K_2 A (\theta - 20)}{0.05} \Rightarrow -20 - \theta = \theta - 20 \quad \therefore \theta = 0^\circ\text{C}$$

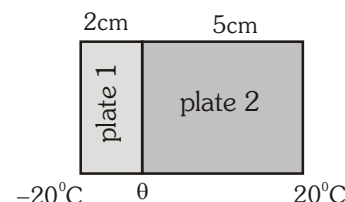


Illustration 25.

Two thin concentric shells made of copper with radius r_1 and r_2 ($r_2 > r_1$) have a material of thermal conductivity K filled between them. The inner and outer spheres are maintained at temperatures T_H and T_C respectively by keeping a heater of power P at the centre of the two spheres. Find the value of P .

Solution : Heat flowing per second through each cross-section of the sphere = $P = i$.

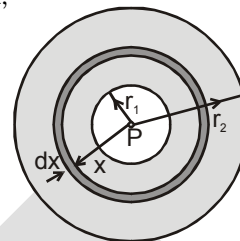
Thermal resistance of the spherical shell of radius x and thickness dx ,

$$dR = \frac{dx}{K \cdot 4\pi x^2}$$

$$\Rightarrow R = \int_{r_1}^{r_2} \frac{dx}{4\pi x^2 \cdot K} = \frac{1}{4\pi K} \left(\frac{1}{r_1} - \frac{1}{r_2} \right)$$

thermal current

$$i = P = \frac{T_H - T_C}{R} = \frac{4\pi K(T_H - T_C)r_1 r_2}{(r_2 - r_1)}$$

**Illustration 26.**

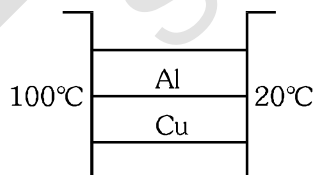
Water in a closed tube is heated with one arm vertically placed above the lamp. In what direction water will begin to circulate along the tube?

Solution

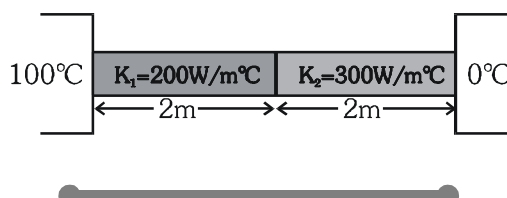
On heating the liquid at A will become lighter and will rise up. This will push the liquid in the tube upwards and so the liquid in the tube will move clockwise i.e. from B to A.

**BEGINNER'S BOX-3**

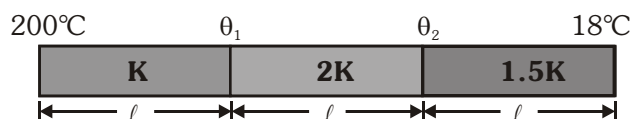
- Explain why :
 - a brass tumbler feels much colder than a wooden tray on a chilly day
 - two layers of cloth of equal thickness provide warmer covering than a single layer of cloth of double thickness ?
 - mud-houses are colder in summer and warmer in winter?
 - In winter birds sit with their wings spread out ?
 - Woollen clothes are warmer than cotton clothes ?
- Two metal cubes with 3 cm-edges of copper and aluminium are arranged as shown in figure. Find



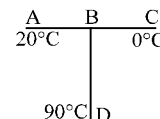
- The total thermal current from one reservoir to the other.
 - The ratio of the thermal current carried by the copper cube to that carried by the aluminium cube. Thermal conductivity of copper is 60 W/m-K and that of aluminium is 40 W/m-K .
- For shown situation, calculate the temperature of the common interface.



4. Calculate θ_1 and θ_2 in shown situation.



5. The temperature at the ends of a uniform rod of length 100 cm are respectively 95°C and 5°C . What will be the temperature at a point 30 cm far from the hotter end ? Also calculate the temperature gradient.
6. Three conducting rods of same material and cross-section are shown in figure. Temperature of A, D and C are maintained at 20°C , 90°C and 0°C . Find the ratio of length BD and BC if there is no heat flow in AB.



3.3 Thermal Radiation

The process of the transfer of heat from one place to another place without heating the intervening medium is called radiation. When a body is heated and placed in vacuum, it loses heat even when there is no medium surrounding it. The heat can not go out from the body by the process of conduction or convection since both of these process require the presence of a material medium between source and surrounding objects. The process by which heat is lost in this case is called radiation. This does not require the presence of any material medium.

It is radiation by which heat from the Sun reaches the Earth. Radiation has the following properties:

- Radiant energy travels in straight lines and when some object is placed in the path, its shadow is formed at the detector.
- It is reflected and refracted or can be made to interfere. The reflection or refraction are exactly as in case of light.
- It can travel through vacuum.
- Intensity of radiation follows the law of inverse square.
- Thermal radiation can be polarised in the same way as light by transmission through a nicol prism.
- Radiation takes place in solid, liquid & gases.

All these and many other properties establish that heat radiation has nearly all the properties possessed by light and these are also electromagnetic waves with the only difference of wavelength or frequency. The wavelength of thermal radiation is larger than that of visible light.

- When radiation passes through any medium then radiations slightly absorbed by medium according to its absorptive power so temperature of medium slightly increases.
- In order to obtain a spectrum of radiation, a special prism is used like KCl prism, Rock salt prism Flourspar prism. Normal glass prism or Quartz prism can not be used (because it absorbs some radiation).
- Radiation intensity is measured with a specific device named as Bolometer.
- Heat radiation are always obtained in infra-red region of electromagnetic wave spectrum so they are called Infra-red rays.
- Thermal radiation when incident on a surface, then exert pressure on the surface which is known as Radiation Pressure.

Basic Fundamental definitions

- Absorptive power or absorptive coefficient (a) :** The ratio of amount of radiation absorbed by a surface

(Q_a) to the amount of radiation incident (Q) upon it, is defined as the coefficient of absorption $a = \frac{Q_a}{Q}$. It

is unitless and dimensionless.

- Spectral absorptive power (a_λ)** $a_\lambda = \frac{Q a_\lambda}{Q_\lambda}$: Also called monochromatic absorptive coefficient

At a given wavelength $a = \int_0^\infty a_\lambda d\lambda$. For ideal black body a_λ and $a = 1$, a and a_λ are unitless

- Emissive power (e)** : The amount of heat radiation emitted by unit surface area in unit second at a particular temperature. **SI UNIT** : $J/m^2\text{-s}$ or watt/m^2
- Spectral Emissive power (e_λ)** : The amount of heat radiation emitted by unit area of the body in one second in unit spectral region at a given wavelength.

Emissive power or total emissive power $e = \int_0^\infty e_\lambda d\lambda$

SI UNIT : $W/m^2\text{-}\text{\AA}$

Emissivity (e)

- Absolute emissivity or emissivity** : Radiation energy given out by a unit surface area of a body in unit time corresponding to unit temperature difference w.r.t. the surroundings is called Emissivity.

SI unit : $W/m^2\text{ K}$

- Relative emissivity (e_r)** : $e_r = \frac{Q_{GB}}{Q_{IBB}} = \frac{e_{GB}}{E_{IBB}} = \frac{\text{emitted radiation by gray body}}{\text{emitted radiation by ideal black body}}$

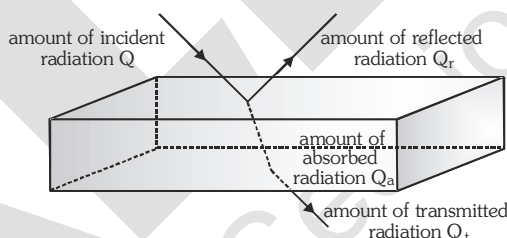
GB = gray or general body, IBB = Ideal black body

- (i) No unit (ii) For ideal black body $e_r = 1$ (iii) range $0 < e_r < 1$

Spectral, emissive, absorptive and transmittive power of a given body surface

Due to incident radiations on the surface of a body following phenomena occur by which the radiation is divided into three parts. (a) Reflection (b) Absorption (c) Transmission

- From energy conservation**



$$Q = Q_r + Q_a + Q_t \Rightarrow \frac{Q_r}{Q} + \frac{Q_a}{Q} + \frac{Q_t}{Q} = 1 \Rightarrow r + a + t = 1$$

$$\text{Reflective Coefficient } r = \frac{Q_r}{Q}, \text{ Absorptive Coefficient } a = \frac{Q_a}{Q},$$

$$\text{Transmittive Coefficient } t = \frac{Q_t}{Q}$$

$$r = 1 \text{ and } a = 0, \quad t = 0 \Rightarrow \text{Perfect reflector}$$

$$a = 1 \text{ and } r = 0, \quad t = 0 \Rightarrow \text{Ideal absorber (ideal black body)}$$

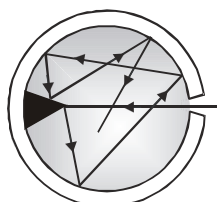
$$t = 1 \text{ and } a = 0, \quad r = 0 \Rightarrow \text{Perfect transmitter (diathermanous)}$$

$$\text{Reflection power (r)} = \left[\frac{Q_r}{Q} \times 100 \right] \%, \text{ Absorption power (a)} = \left[\frac{Q_a}{Q} \times 100 \right] \%$$

$$\text{Transmission power (t)} = \left[\frac{Q_t}{Q} \times 100 \right] \%$$

Ideal Black Body

- A body surface which absorbs all incident thermal radiations at low temperature, irrespective of their wave length and emits out all these absorbed radiations at high temperature is assumed to be an ideal black body surface.
- The identical parameters of an ideal black body is given by $a = a_\lambda = 1$ and $r = 0 = t$, $e_r = 1$



Ferry's ideal black body

- The nature of emitted radiations from surface of ideal black body only depends on its temperature
- The radiations emitted from surface of ideal black body are called as either full or white radiations.
- At any temperature the spectral energy distribution curve for surface of an ideal black body is always continuous and according to this concept if the spectrum of a heat source obtained is continuous then it must be ideal black body like kerosene lamp; oil lamp, heating filament etc.
- There are two experimentally ideal black body
(a) Ferry's ideal black body (b) Wien's ideal black body.
- At low temperature, surface of ideal black body is a perfect absorber and at a high temperature it proves to be a perfect emitter.
- An ideal black body need not be of black colour (eg. Sun).

Prevost's theory of heat energy exchange

According to Prevost, at every possible temperature (except zero kelvin temperature) there is a continuous heat energy exchange between a body and its surrounding and this exchange carry on for infinite time.

The relation between temperature difference of body with its surrounding decides whether the body experience cooling effect or heating effect.

When a cold body is placed in the hot surrounding : The body radiates less energy and absorbs more energy from the surrounding, therefore the temperature of body increases. (Heating effect)

When a hot body placed in cooler surrounding : The body radiates more energy and absorbs less energy from the surroundings. Therefore temperature of body decreases. (cooling effect)

When the temperature of a body is equal to the temperature of the surrounding

The energy radiated per unit time by the body is equal to the energy absorbed per unit time by the body, therefore its temperature remains constant and the body is in thermal equilibrium with surrounding. Hence no heating and cooling effects are seen.

KIRCHHOFF'S LAW

At a given temperature for all bodies the ratio of their spectral emissive power (e_λ) to spectral absorptive power (a_λ) is constant and this constant is equal to spectral emissive power (E_λ) of the ideal black body at same temperature

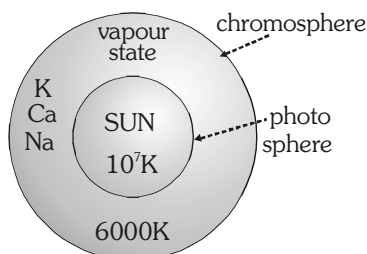
$$\frac{e_\lambda}{a_\lambda} = E_\lambda = \text{constant} \quad \left[\frac{e_\lambda}{a_\lambda} \right]_1 = \left[\frac{e_\lambda}{a_\lambda} \right]_2 = \text{constant} \quad \text{Hence} \quad \boxed{e_\lambda \propto a_\lambda}$$

Good absorbers are good emitters and bad absorbers are bad emitters

Applications of Kirchoff Law

• Fraunhofer's lines

Fraunhofer lines are dark lines in the spectrum of the Sun. When white light emitted from the central core of the Sun (Photosphere) passes through its atmosphere (chromosphere) some of the radiations are absorbed by the gases present, resulting in dark lines in the spectrum of Sun. At the time of total solar eclipse direct light rays emitted from photosphere cannot reach on the Earth and only rays from chromosphere are able to reach on the Earth surface. At that time we observe bright fraunhofer lines.



• In deserts days are hot and nights cold

Sand is rough and black, so it is a good absorber and hence in deserts, days (when radiation from Sun is incident on sand) will be very hot. Now in accordance with Kirchhoff's Law, good absorber is a good emitter. So nights (when sand emits radiation) will be cold.

Stefan's Law

The amount of radiation emitted per second per unit area by ideal black body is directly proportional to the fourth power of its absolute temperature.

Amount of radiation emitted $E \propto T^4$ where T = temperature of ideal black body (in K)
 $E = \sigma T^4$ (This law is true for only ideal black body)

SI Unit : $E = \text{watt/m}^2$ $\sigma \rightarrow$ Stefan's constant = $5.67 \times 10^{-8} \text{ watt /m}^2 \text{ K}^4$ (universal constant)

Dimensions of σ : $M^1 L^0 T^{-3} \theta^{-4}$

Total radiation energy emitted out by surface of area A in time t :

Ideal black body $Q_{\text{IBB}} = \sigma A T^4 t$ and for any other body $Q_{\text{GB}} = e_r \sigma A T^4 t$

Rate of emission of radiation

When temperature of surrounding T_0 (Let $T_0 < T$)

Rate of emission of radiation from per unit area of ideal black body surface $E_1 = \sigma T^4$

Rate of emission or absorption of radiation (per unit area) from surrounding $E_2 = \sigma T_0^4$

Net rate of loss of radiation per unit area from ideal black body surface is

$$E = E_1 - E_2 = \sigma T^4 - \sigma T_0^4 = \sigma (T^4 - T_0^4)$$

Net loss of radiation energy from entire surface area in time t is $Q_{\text{IBB}} = \sigma A (T^4 - T_0^4) t$

For any other body $Q_{\text{GB}} = e_r A \sigma (T^4 - T_0^4) t$

If in time dt the net heat energy loss for ideal black body is dQ and because of this its temperature falls by dT

Rate of loss of heat (IBB)
$$R_H = \frac{dQ}{dt} = \sigma A (T^4 - T_0^4) \text{ J/s}$$

It is also equal to emitted power or radiation emitted per second

Rate of fall in temperature (Rate of cooling)
$$R_F = \frac{dT}{dt} = \frac{\sigma A}{ms} (T^4 - T_0^4) \left[\because \frac{dQ}{dt} = ms \frac{dT}{dt} \right]$$

Body	R_H	R_F
(i) Two solid sphere (same material) (same T, T_0, s, ρ) (different radius r_1, r_2)	$R_H \propto r^2$	$R_F \propto \frac{1}{r}$
(ii) Two solid sphere (different material) (same T, T_0)	$R_H \propto r^2$	$R_F \propto \frac{1}{rps}$
(iii) Different shape bodies (Cube, sphere, cylinder flat surface) (const. T, T_0, V , same materials) surface	$R_H \propto A$ • maximum for flat surface • minimum for spherical	$R_F \propto \frac{A}{V}$ • maximum for flat surface • minimum for spherical surface
(iv) Two sphere (one solid and another hollow) (T, T_0, s, A are same)	• R_H is same for both	R_F • maximum for hollow sphere • minimum for solid sphere
(v) Different bodies made of different material (T, T_0, M, A)(s-different)	• R_H is same for all bodies	$R_F \propto \frac{1}{s}$

Where T and T_0 absolute temperature of body and surrounding, M = mass of body,
 s = Specific heat, ρ = density, V = volume.

• **When a body cools by radiation, its cooling depends on :**

- Nature of radiating surface : greater the emissivity (e_r), faster will be the cooling.
- Area of radiating surface : greater the area of radiating surface, faster will be the cooling.
- Mass of radiating body : greater the mass of radiating body slower will be the cooling.
- Specific heat of radiating body : greater the specific heat of radiating body slower will be the cooling.
- Temperature of radiating body : greater the temperature of radiating body faster will be the cooling.

NEWTON'S LAW OF COOLING

Rate of cooling $\left(\frac{dT}{dt}\right)$ is directly proportional to excess of temperature of the body over that of surrounding.

[(when $(T - T_0) \neq 35^\circ\text{C}$) $-\frac{dT}{dt} \propto (T - T_0)$]

T = temperature of body [all temperatures in $^\circ\text{C}$]

T_0 = temperature of surrounding,

$T - T_0$ = excess of temperature ($T > T_0$)

If the temperature of body is decreased by dT in time dt then rate of fall of temperature $-\frac{dT}{dt} = K(T - T_0)$

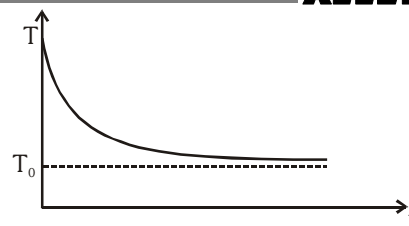
Where negative sign indicates that the rate of cooling is decreasing with time.

For Numerical Problems, Newton's Law of cooling

If the temperature of body decreases from T_1 to T_2 and temperature of surroundings is T_0 then

$$\text{average excess of temperature} = \left[\frac{T_1 + T_2}{2} - T_0 \right]$$

$$\Rightarrow \left[\frac{T_1 - T_2}{t} \right] = +K \left[\frac{T_1 + T_2}{2} - T_0 \right]$$

**Limitations of Newton's Law of Cooling**

- Temperature difference should not exceed 35°C , $(T - T_0) \propto 35^\circ \text{C}$
- Loss of heat should only be by radiation.
- This law is an extended form of Stefan-Boltzman's law.

Derivation of Newton's law from Stefan's Boltzman law

$$\frac{dT}{dt} = \frac{\sigma A}{ms} (T^4 - T_0^4) \quad \begin{cases} T - T_0 = \Delta T \\ T = T_0 + \Delta T \end{cases}$$

$$\frac{dT}{dt} = \frac{\sigma A}{ms} [(T_0 + \Delta T)^4 - T_0^4] \quad \text{If } x \ll 1 \text{ then } (1 + x)^n = 1 + nx \text{ (Binomial theorem)}$$

$$\frac{dT}{dt} = \frac{\sigma A}{ms} \left[T_0^4 \left(1 + \frac{\Delta T}{T_0} \right)^4 - T_0^4 \right] = \frac{\sigma A}{ms} T_0^4 \left[\left(1 + \frac{\Delta T}{T_0} \right)^4 - 1 \right] = \frac{\sigma A}{ms} T_0^4 \left[1 + 4 \frac{\Delta T}{T_0} - 1 \right]$$

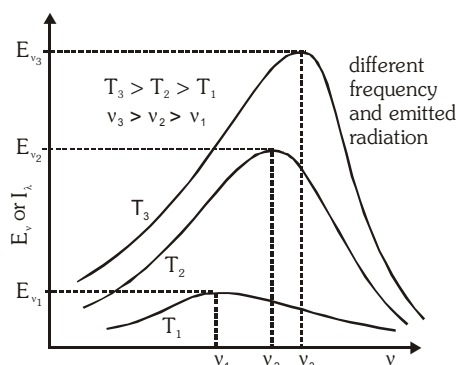
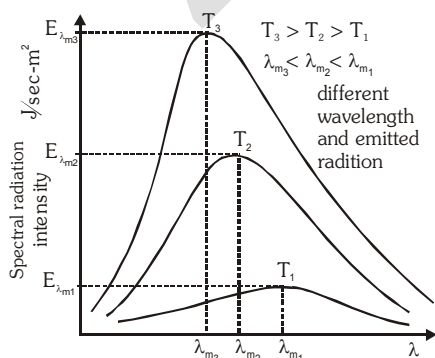
$$\frac{dT}{dt} = \left[4 \frac{\sigma A}{ms} T_0^3 \right] \Delta T \Rightarrow \frac{dT}{dt} = K \Delta T \quad \text{here constant } K = \frac{4\sigma A T_0^3}{ms}$$

$$\text{Newton's law of cooling} \quad \frac{dT}{dt} \propto \Delta T \quad (\text{for small temperature difference})$$

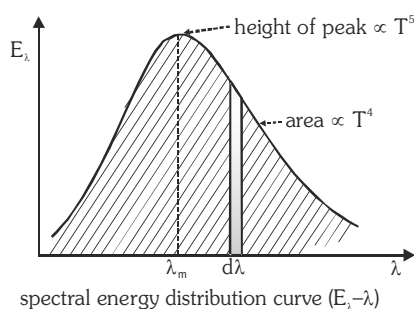
Spectral Energy distribution curve of Black Body radiations

Practically given by : Lumers and Pringshem

Mathematically given by : Plank



Results from these Graphs



(i) $\lambda_m \propto \frac{1}{T}$ Wien's displacement law

$$\lambda_m T = b$$

$$\lambda_{m_1} T_1 = \lambda_{m_2} T_2$$

(ii) $E_{\lambda_m} \propto T^5$

(iii) i.e. Area $\int_0^\infty E_\lambda d\lambda = E = \sigma T^4$

Hence $\frac{A_1}{A_2} = \left[\frac{T_1}{T_2} \right]^4$

Area between curve and λ axis gives the emissive power of body

Wein's Displacement Law

The wavelength corresponding to maximum emission of radiation decrease with increasing temperature $\left[\lambda_m \propto \frac{1}{T} \right]$. This is known as Wein's displacement law.

$$\lambda_m T = b$$

where b is Wein's constant = 2.89×10^{-3} mK.

Dimensions of b : $M^0 L^1 T^0 \theta^1$

Relation between frequency and temperature $\nu_m = \frac{c}{b} T$ [$c = \lambda \times \nu$]

Solar constant 'S'

The Sun emits radiant energy continuously in space of which an in significant part reaches the Earth. The solar radiant energy received per unit area per unit time by a black surface held at right angles to the Sun's rays and placed at the mean distance of the Earth (in the absence of atmosphere) is called solar constant. The solar constant S is taken to be 1340 watt/m^2 or $1.937 \text{ Cal/cm}^2\text{-minute}$

• Temperature of the Sun

Let R be the radius of the Sun and ' d ' be the radius of Earth's orbit around the Sun. Let E be the energy emitted by the Sun per second per unit area. The total energy emitted by the Sun in one second = $E \cdot A = E \times 4\pi R^2$. (This energy is falling on a sphere of radius equal to the radius of the Earth's orbit around the Sun i.e., on a sphere of surface area $4\pi d^2$)

So, The energy falling per unit time per unit area of Earth

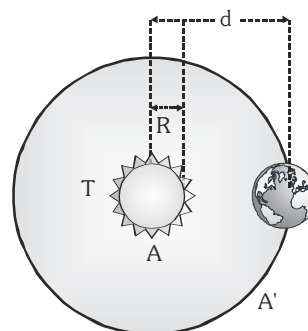
$$= \frac{4\pi R^2 \times E}{4\pi d^2} = \frac{E R^2}{d^2}$$

$$R = 7 \times 10^8 \text{ m}, \quad d = 1.5 \times 10^{11} \text{ m}, \quad \sigma = 5.7 \times 10^{-8} \text{ W m}^{-2} \text{ K}^{-4}$$

Solar constant $S = \frac{E R^2}{d^2}$

By Stefan's Law $E = \sigma T^4$

$$S = \frac{\sigma T^4 R^2}{d^2} \Rightarrow T = \left[\frac{S \times d^2}{\sigma \times R^2} \right]^{\frac{1}{4}} = \left[\frac{1340 \times (1.5 \times 10^{11})^2}{5.7 \times 10^{-8} \times (7 \times 10^8)^2} \right]^{\frac{1}{4}} = 5732 \text{ K}$$



GOLDEN KEY POINTS

- At absolute zero temperature (zero kelvin) all atoms of a given substance remains in ground state, so, at this temperature emission and absorption of radiation from any substance is impossible, so Prevost's heat energy exchange theory does not applied at this temperature, so it is called **limiting temperature** of prevost's theory.
- With the help of Prevost's theory rate of cooling of any body w.r.t. its surroundings can be worked out (applied to Stefan Boltzman law, Newton's law of cooling.)
- For a constant temperature the spectral emissive power of an ideal black body is a constant parameter
- The practical confirmation of Kirchhoff's law carried out by Rishi apparatus and the main base of this apparatus is a Lesslie container.
- The main conclusion predicted from Kirchhof's law can be expressed as
 Good absorber \Rightarrow Good emitter
 Bad absorber \Rightarrow Bad emitter
 (at Low temperature) (at high temperature)
- If all of T, T_0, m, s, V, ρ , are same for different shape body then R_F and R_H will be maximum for the flat surface.
- If a solid and hollow sphere are taken with all the parameters same then hollow will cool down at fast rate.
- Rate of temperature fall, $R_F \propto \frac{1}{s} \propto \frac{dT}{dt}$
 so $dt \propto s$. If condition of specific heat is $s_1 > s_2 > s_3$
 and if all cooled same temperature i.e. temperature fall is also identical for all then required time $t \propto s$
 $\therefore t_1 > t_2 > t_3$
- Spectral energy distribution curves are continuous. At any temperature in all possible wavelength radiation between $(0 - \infty)$ are emitted but quantity of radiations are different for different wavelength.
- As the wave length increases, the amount of radiation emitted first increase, becomes maximum and then decreases.
- At a particular temperature the area enclosed between the spectral energy curve and wavelength axis shows the emissive power of the body.

$$\text{Area} = \int_0^{\infty} E_{\lambda} d\lambda = E = \sigma T^4$$

Illustrations

Illustration 27.

If λ_m for the moon is 14.5 micron, then find its temperature.

Solution

Wien's displacement law $\lambda_m T = b$

$$\therefore T = \frac{b}{\lambda_m} = \frac{2.89 \times 10^{-3}}{14.5 \times 10^{-6}} = 199.3K$$

Illustration 28.

Total radiation incident on body is 400 J, If 20% of incident radiation reflected back and 120 J is absorbed by body. Then find out transmittive power in percentage.

Solution

$$Q = Q_t + Q_r + Q_a \Rightarrow 400 = 80 + 120 + Q_t \Rightarrow Q_t = 200$$

$$\text{So transmittive power is } \frac{Q_t}{Q} \times 100\% = 50\%$$

Illustration 29.

The operating temperature of a tungsten filament in an incandescent lamp is 2000 K and its emissivity is 0.3. Find the surface area of the filament of a 25 watt lamp. Stefan's constant $\sigma = 5.67 \times 10^{-8} \text{ Wm}^{-2} \text{ K}^{-4}$

Solution

\therefore Rate of emission = wattage of the lamp

$$\therefore W = Ae\sigma T^4 \Rightarrow A = \frac{W}{e\sigma T^4} = \frac{25}{0.3 \times 5.67 \times 10^{-8} \times (2000)^4} = 0.918 \times 10^{-4} \text{ m}^2$$

Illustration 30.

Draw a graph between $\log E$ and $\log T$

Solution

$$E = \sigma T^4 \quad (\text{taking log})$$

$$\log E = \log (\sigma T^4)$$

$$\log E = 4\log T + \log \sigma \quad \text{This is equivalent to } y = mx - C \quad (\sigma < 1 \text{ so its log is a negative quantity})$$

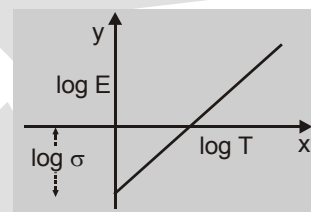


Illustration 31.

If temperature of ideal black body is increased by 50%, what will be percentage increase in quantity of radiations emitted from its surface.

Solution

$$E \propto T^4 \quad \text{and} \quad \therefore E' \propto (1.5)^4 T^4 \propto \left[\frac{15}{10}\right]^4 T^4 \propto \left[\frac{3}{2}\right]^4 T^4 \propto \frac{81}{16} T^4$$

$$\frac{E' - E}{E} \times 100\% = \left[\frac{\frac{81}{16} T^4 - T^4}{T^4} \right] \times 100\% = 406\% \approx 400\%$$

Illustration 32.

If temperature of ideal black body is decreased from T to $\frac{T}{2}$ then find out percentage loss in emissive rate

Solution

$$E \propto T^4, \quad E' \propto \left[\frac{T}{2}\right]^4 = \frac{T^4}{16}$$

$$\left[\frac{E - E'}{E} \right] \times 100\% = \left[1 - \frac{1}{16} \right] \times 100\% = \frac{15}{16} \times 100\% \approx 94\% \quad \text{Remaining is } 6\% \text{ (Approx.)}$$

Illustration 33.

Calculate the temperature at which a perfect black body radiates at the rate of 5.67 W cm^{-2} . Stefan's constant is $5.67 \times 10^{-8} \text{ J s}^{-1} \text{ m}^{-2} \text{ K}^{-4}$.

Solution

Given $E = 5.67 \text{ W cm}^{-2} = 5.67 \times 10^4 \text{ W m}^{-2}$, $\sigma = 5.67 \times 10^{-8} \text{ J s}^{-1} \text{ m}^{-2} \text{ K}^{-4}$

$$\text{Using, } E = \sigma T^4; \quad T^4 = \frac{E}{\sigma} \quad \text{or } T = \left[\frac{E}{\sigma} \right]^{\frac{1}{4}} = \left[\frac{5.67 \times 10^4}{5.67 \times 10^{-8}} \right]^{\frac{1}{4}} = (10^{12})^{\frac{1}{4}} = 1000 \text{ K}$$

Illustration 34.

The temperature of furnace is 2000°C , in its spectrum the maximum intensity is obtained at about 4000\AA . If the maximum intensity is at 2000\AA calculate the temperature of the furnace in $^\circ\text{C}$.

Solution

by using $\lambda_m T = b$, $4000 (2000+273) = 2000(T) \Rightarrow T = 4546\text{K}$

The temperature of furnace = $4546 - 273 = 4273^\circ\text{C}$

Illustration 35.

Two bodies A and B have thermal emissivities of 0.01 and 0.81 respectively. The outer surface areas of the two bodies are same, the two bodies emit total radiant power at the same rate. The wavelength λ_B corresponding to maximum spectral radiance of B is shifted from the wavelength corresponding to maximum spectral radiance in the radiation of A by $1.0 \mu\text{m}$. If the temperature of A is 5802K , Calculate :-

- The temperature of B
- Wavelength λ_B

Solution

- As both bodies A and B having same radiant power

$$\therefore P_A = P_B \Rightarrow e_A \sigma A_A T_A^4 = e_B \sigma A_B T_B^4 \Rightarrow (0.01) \sigma A T_A^4 = (0.81) \sigma A T_B^4$$

$$T_B = \left(\frac{0.01}{0.81} \right)^{\frac{1}{4}} T_A = \frac{T_A}{3} = \frac{5802}{3} = 1934 \text{ K}$$

- According to wein's displacement law

$$\lambda_A T_A = \lambda_B T_B \Rightarrow \lambda_B = \left(\frac{5802}{1934} \right) \lambda_A = 3\lambda_A$$

$$\text{As } \lambda_B - \lambda_A = 1 \mu\text{m} \Rightarrow \lambda_B - \frac{\lambda_B}{3} = 1 \mu\text{m} \Rightarrow \frac{2\lambda_B}{3} = 1 \mu\text{m} \Rightarrow \lambda_B = 1.5 \mu\text{m}$$

Illustration 36.

When a metallic body is heated in a furnace, then what colour will appear as temperature increases.

Solution. $T \propto \frac{1}{\lambda_m}$

As Temperature increases colour of body will appear from red, yellow, green, blue and then white.

Illustration 37.

Define (i) Steady state and (ii) Temperature gradient in conduction of heat through a conducting rod.

Solution

- When one end of a rod is heated, the temperature of various points of the rod changes continuously but after some time a state is reached, when the temperature of each cross-section becomes steady which is called steady state. In this state the heat received by any section will be totally transferred to the next section so no heat is absorbed by any cross section.
- Temperature gradient is defined as the rate of change of temperature with distance in the direction of flow of heat.

Illustration 38.

Assuming Newton's law of cooling to be valid. The temperature of body changes from 60°C to 40°C in 7 minutes. Temperature of surroundings being 10°C , Find its temperature after next 7 minutes.

Solution

According to Newton's law of cooling $\frac{T_1 - T_2}{t} = K \left(\frac{T_1 + T_2}{2} - T_0 \right)$

Since the temperature decreases from 60°C to 40°C in 7 minutes

$$\frac{60 - 40}{7} = K \left(\frac{60 + 40}{2} - 10 \right) \Rightarrow \frac{20}{7} = K (50 - 10) \Rightarrow K = \frac{1}{14}$$

If the temperature of object becomes T_2 in next 7 minutes then $\frac{40 - T_2}{7} = \frac{1}{14} \left(\frac{40 + T_2}{2} - 10 \right)$

$$\Rightarrow 40 - T_2 = \frac{1}{4} (40 + T_2 - 20) \Rightarrow 160 - 4T_2 = 20 + T_2$$

$$\Rightarrow 5T_2 = 140 \Rightarrow T_2 = 28^{\circ}\text{C}$$

BEGINNER'S BOX-4

1. Explain why :

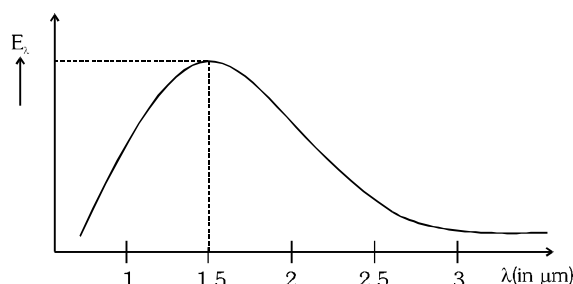
- a body with large reflectivity is a poor emitter.
- an optical pyrometer (for measuring high temperatures) calibrated for an ideal black body radiation gives too low value for the temperature of a red hot iron piece in the open, but gives a correct value for the temperature when the same piece is in the furnace
- the earth without its atmosphere would be inhospitably cold
- Heat is generated continuously in an electric heater but its temperature remains constant after some time.
- The bottom of a cooking vessel is made black ?
- On winter night you feel warmer when clouds cover the sky than when the sky is clear.
- A thermos or vacuum flask can keep hot things hot and cold things cold for a long time, how?

2. Two spherical ideal black bodies of radii r_1 and r_2 are having surface temperature T_1 and T_2 respectively, If both radiate the same power. Then calculate the ratio of $\frac{T_1}{T_2}$.

3. If a liquid takes 30 sec. in cooling from 80°C to 70°C and 70 sec in cooling from 60°C to 50°C , then find the room temperature.

4. A body cools in 10 minutes from 60°C to 40°C . What will be its temperature after next 10 minutes? The temperature of the surrounding is 10°C .

5. Calculate the temperature of the black body from given graph.



KINETIC THEORY OF GASES

The properties of the gases are entirely different from those of solid and liquid. In case of gases, thermal expansion is very large as compared to solids and liquids. To state the conditions of a gas, its volume, pressure and temperature must be specified.

Intermolecular force Solid > liquid > real gas > ideal gas (zero)

Potential energy Solid < liquid < real gas < ideal gas (zero)

At a given temperature for solid, liquid and gas:

- (i) Internal kinetic energy : Same for all
- (ii) Internal potential Energy : Maximum for ideal gas (PE = 0) and Minimum for solids (PE = -ve)
- (iii) Internal Energy : Maximum for Ideal gas and Minimum for solid

Notation :

μ = Molar amount of gas = $M/M_w = N/N_0$

N_0 = Avogadro constant = 6.023×10^{23} molecules/mole

R = Universal gas constant = $8.31 \text{ J/mole}^\circ\text{K}$

$\approx 2 \text{ cal/mole}^\circ\text{K}$, [Dimension = $\text{ML}^2\text{T}^{-2}\mu^{-1}\theta^{-1}$]

r = Specific gas constant ($r = R/M_w$)

k = Boltzmann constant ($k = R/N_0$) = $1.38 \times 10^{-23} \text{ J/kelvin}$

m = mass of a gas molecule

M = mass of gas

M_w = molecular weight

N = No. of gas molecules

n = Molecular density ($n = N/V$)

ρ = gas density ($\rho = M/V = mn$)

N.T.P. (Normal temperature and pressure)		S.T.P. (Standard Temperature and Pressure)
Temperature	$0^\circ\text{C} = 273.15 \text{ K}$	$0.01^\circ\text{C} = 273.16 \text{ K}$
Pressure	$1 \text{ atm} = 1.01325 \times 10^5 \text{ N/m}^2$ $= 1.01325 \times 10^5 \text{ pascal}$	1 atm
Volume	22.4 litre	22.4 litre

$1 \text{ lit} = 10^{-3} \text{ m}^3 = 10^3 \text{ cm}^3 = 10^3 \text{ cc} = 10^3 \text{ ml}$

$\rho_{\text{water}} = 1000 \text{ kg/m}^3 = 1 \text{ g/cc}$, $\rho_{\text{ice}} = 900 \text{ kg/m}^3 = 0.9 \text{ g/cc}$

Ideal Gas Concept

- Volume of gas molecules is negligible as compared to volume of container.
So volume of gas = volume of container (Except 0 K)
- No intermolecular forces act between gas molecules.

Properties of Ideal Gas

- A gas which follows all gas laws and gas equation at every possible temperature and pressure is known as ideal or perfect gas.
- Ideal gas molecules can do only translational motion, so their kinetic energy is only translational kinetic energy
- Ideal gas can not be liquified because IMF is zero
- Potential energy of ideal gas is zero so internal energy of ideal gas is perfectly translational K.E. of gas. It is directly proportional to absolute temperature.

So, internal energy depends only and only on its temperature.

$$E_{\text{trans}} \propto T$$

For a substance $U = U_{\text{KE}} + U_{\text{PE}}$

U_{KE} : depends only on T , U_{PE} : depends upon intermolecular forces (Always negative)

- Specific heat of ideal gas is constant quantity and it does not change with temperature
- All real gases behaves as an ideal gas at high temperature and low pressure and low density.
- Gas molecules have point mass and negligible volume and velocity is very high (10^7 cm/s). That's why there is no effect of gravity on them.

Equation of state for Ideal gas

$$PV = \mu RT \Rightarrow PV = \frac{M}{M_w} RT = \left[\frac{mN}{mN_0} \right] RT = \left[\frac{R}{N_0} \right] N T = NkT \quad \text{where } \mu = \text{number of moles of gas}$$

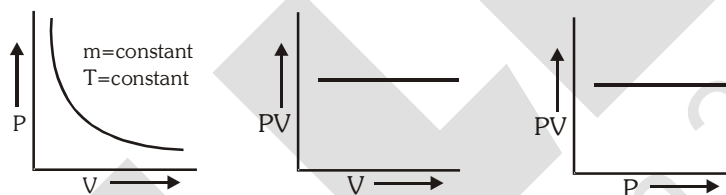
k = Boltzmann constant

$$\Rightarrow \frac{P}{\rho} = \frac{RT}{M_w} = \frac{kT}{m}$$

Gas Laws

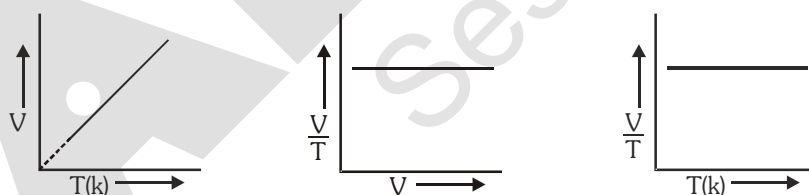
Boyle's Law

According to it for a given mass of an ideal gas at constant temperature, the volume of a gas is inversely proportional to its pressure, i.e., $V \propto \frac{1}{P}$ if m and T = Constant $\Rightarrow P_1 V_1 = P_2 V_2$



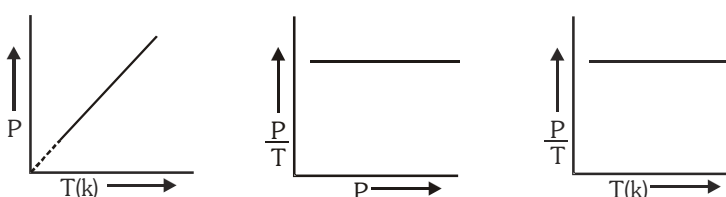
Charles's Law

According to it for a given mass of an ideal gas at constant pressure, volume of a gas is directly proportional to its absolute temperature, i.e., $V \propto T$ if m and P = Constant $\Rightarrow \frac{V_1}{V_2} = \frac{T_1}{T_2}$



Gay-Lussac's Law

According to it, for a given mass of an ideal gas at constant volume, pressure of a gas is directly proportional to its absolute temperature, i.e., $P \propto T$ if m and V = constant $\Rightarrow \frac{P_1}{P_2} = \frac{T_1}{T_2}$



Avogadro's Law

According to it, at same temperature and pressure, equal volume of all gases contain equal number of molecules, i.e., $N_1 = N_2$ if P, V and T are same.

Dalton's Partial Pressure Mixture Law :

According to it, the pressure exerted by mixture of non-reactive gases is equal to the sum of partial pressure of each component gases present in the mixture, i.e., $P = P_1 + P_2 + \dots$

Illustrations

Illustration 39.

By increasing temperature of gas by 5°C its pressure increases by 0.5% from its initial value at constant volume then what is initial temperature of gas ?

Solution

$$\therefore \text{At constant volume } T \propto P \quad \therefore \frac{\Delta T}{T} \times 100 = \frac{\Delta P}{P} \times 100 = 0.5 \Rightarrow T = \frac{5 \times 100}{0.5} = 1000 \text{K}$$

Illustration 40.

Calculate the value of universal gas constant at STP.

Solution

Universal gas constant is given by $R = \frac{PV}{T}$

One mole of all gases at S.T.P. occupy volume $V = 22.4 \text{ litre} = 22.4 \times 10^{-3} \text{ m}^3$

$P = 760 \text{ mm of Hg} = 760 \times 10^{-3} \times 13.6 \times 10^3 \times 9.80 \text{ N m}^{-2}$, $T = 273.16 \text{ K}$

$$\therefore R = \frac{760 \times 10^{-3} \times 13.6 \times 10^3 \times 9.80 \times 22.4 \times 10^{-3}}{273.16} = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$$

Illustration 41.

A closed container of volume 0.02 m^3 contains a mixture of neon and argon gases at a temperature of 27°C and pressure of $1 \times 10^5 \text{ N/m}^2$. The total mass of the mixture is 28 g . If the gram molecular weights of neon and argon are 20 and 40 respectively, find the mass of the individual gases in the container, assuming them to be ideal. Given : $R = 8.314 \text{ J/mol-K}$.

Solution

Let m gram be the mass of neon. Then, the mass of argon is $(28 - m)\text{g}$.

$$\text{Total number of moles of the mixture, } \mu = \frac{m}{20} + \frac{28 - m}{40} = \frac{28 + m}{40} \quad \dots(i)$$

$$\text{Now, } \mu = \frac{PV}{RT} = \frac{1 \times 10^5 \times 0.02}{8.314 \times 300} = 0.8 \quad \dots(ii)$$

$$\text{By (i) and (ii), } \frac{28 + m}{40} = 0.8 \Rightarrow 28 + m = 32 \Rightarrow m = 4 \text{ gram}$$

$$\text{or mass of argon} = (28 - 4)\text{g} = 24 \text{ g}$$

Illustration 42.

Calculate the temperature of the Sun if density is 1.4 g cm^{-3} , pressure is 1.4×10^9 atmosphere and average molecular weight of gases in the Sun is 2 g/mole . [Given $R = 8.4 \text{ J mol}^{-1}\text{K}^{-1}$]

Solution

$$PV = \mu RT \Rightarrow T = \frac{PV}{\mu R} \dots (i) \quad \text{But } \mu = \frac{M}{M_w} \text{ and } \rho = \frac{M}{V} \therefore \mu = \frac{\rho V}{M_w}$$

$$\text{From equation (i)} \quad T = \frac{PVM_w}{\rho VR} = \frac{PM_w}{\rho R} = \frac{1.4 \times 10^9 \times 1.01 \times 10^5 \times 2 \times 10^{-3}}{1.4 \times 1000 \times 8.4} = 2.4 \times 10^7 \text{ K}$$

Illustration 43.

At the top of a mountain a thermometer reads 7°C and barometer reads 70 cm of Hg . At the bottom of the mountain they read 27°C and 76 cm of Hg respectively. Compare the density of the air at the top with that at the bottom.

Solution

$$\text{By gas equation } PV = \frac{M}{M_w} RT \Rightarrow \frac{P}{\rho T} = \frac{R}{M_w} \left[\because \mu = \frac{M}{M_w} \text{ and } \frac{M}{V} = \rho \right]$$

$$\text{Now as } M_w \text{ and } R \text{ are same for top and bottom } \left[\frac{P}{\rho T} \right]_T = \left[\frac{P}{\rho T} \right]_B \text{ So } \frac{\rho_T}{\rho_B} = \frac{P_T}{P_B} \times \frac{T_B}{T_T} = \frac{70}{76} \times \frac{300}{280} = \frac{75}{76} = 0.9868$$

Illustration 44.

A sample of oxygen of volume of 500 cc at a pressure of 2 atm is compressed to a volume of 400 cc . What pressure is needed to do this if the temperature is kept constant?

Solution

$$\text{Temperature is constant, so } P_1 V_1 = P_2 V_2 \therefore P_2 = P_1 \frac{V_1}{V_2} = 2 \left[\frac{500}{400} \right] = 2.5 \text{ atm}$$

Illustration 45.

A vessel of volume $8.0 \times 10^{-3} \text{ m}^3$ contains an ideal gas at 300 K and 200 kPa . The gas is allowed to leak till the pressure falls to 125 kPa . Calculate the amount of the gas leaked assuming that the temperature remains constant.

Solution

As the gas leaks out, the volume and the temperature of the remaining gas do not change. The number of

moles of the gas in the vessel is given by $\mu = \frac{PV}{RT}$.

The number of moles in the vessel before the leakage is $\mu_1 = \frac{P_1 V}{RT}$ and that after the leakage is $\mu_2 = \frac{P_2 V}{RT}$.

$$\text{The amount leaked is } \mu_1 - \mu_2 = \frac{(P_1 - P_2)V}{RT} = \frac{(200 - 125) \times 10^3 \times 8.0 \times 10^{-3}}{8.3 \times 300} = 0.24 \text{ mole}$$

Illustration 46.

1500 ml of a gas at a room temperature of 23°C is inhaled by a person whose body temperature is 37°C , if the pressure and mass stay constant, what will be the volume of the gas in the lungs of the person ?

Solution

$T_1 = 273 + 23 = 296 \text{ K}$; $T_2 = 273 + 37 = 310 \text{ K}$. Pressure and amount of the gas are kept constant,

$$\text{So } \frac{V_1}{T_1} = \frac{V_2}{T_2} \therefore V_2 = V_1 \times \frac{T_2}{T_1} = 1500 \times \frac{310}{296} = 1570.95 \text{ ml}$$

Illustration 47.

A sample of O_2 is at a pressure of 1 atm when the volume is 100 ml and its temperature is 27°C . What will be the temperature of the gas if the pressure becomes 2 atm and volume remains 100 ml.

Solution

$$T_1 = 273 + 27 = 300 \text{ K}$$

$$\text{For constant volume } \frac{P_1}{T_1} = \frac{P_2}{T_2} \Rightarrow T_2 = T_1 \times \frac{P_2}{P_1} = 300 \times \frac{2}{1} = 600 \text{ K} = 600 - 273 = 327^{\circ}\text{C}$$

BEGINNER'S BOX-5

1. A vessel of volume $8.3 \times 10^{-3} \text{ m}^3$ contains an ideal gas at temperature 27°C and pressure 200 kPa. The gas is allowed to leak till the pressure falls to 100 kPa and temperature increases to 327°C . What is the amount of gas in moles will be leaked out ?
2. Two closed vessels of equal volume contain air at 105 kPa, 300 K and are connected through a narrow tube. If one of the vessels is now maintained at 300 K and other at 400 K, What will be the pressure in the vessels ?
3. The volume of a gas is 1 litre at the pressure $1.2 \times 10^7 \text{ Nm}^{-2}$ and temperature 400 K. Calculate the number of molecules in the gas.
4. A vessel contains two non-reactive gases : neon (monatomic) and oxygen (diatomic) The ratio of their partial pressures is 3:2. Estimate the ratio of
 - (a) number of molecules and
 - (b) mass density of neon and oxygen in the vessel.

Atomic mass of Ne = 20.2 u. molecular mass of $\text{O}_2 = 32.0 \text{ u}$.

The kinetic theory of gases

- Rudolph Clausius (1822–88) and James Clark Maxwell (1831–75) developed the kinetic theory of gases in order to explain gas laws in terms of the motion of the gas molecules. The theory is based on following assumptions regarding the motion of molecules and the nature of the gases.

Basic postulates of Kinetic theory of gases

- Every gas consists of extremely small particles known as molecules. The molecules of a given gas are all identical but are different than those of another gas.
- The molecules of a gas are identical, spherical, rigid and perfectly elastic point masses.
- The size is negligible in comparison to inter molecular distance (10^{-9} m)

Assumptions regarding motion :

- Molecules of a gas keep on moving randomly in all possible direction with all possible velocities.
- The speed of gas molecules lie between zero and infinity (very high speed).
- The number of molecules moving with most probable speed is maximum.

Assumptions regarding collision:

- The gas molecules keep colliding among themselves as well as with the walls of containing vessel. These collision are perfectly elastic. (ie., the total energy before collision = total energy after the collision.)

Assumptions regarding force:

- No attractive or repulsive force acts between gas molecules.
- Gravitational attraction among the molecules is ineffective due to extremely small masses and very high speed of molecules.

Assumptions regarding pressure:

- Molecules constantly collide with the walls of container due to which their momentum changes. This change in momentum is transferred to the walls of the container. Consequently pressure is exerted by gas molecules on the walls of container.

Assumptions regarding density:

- The density of gas is constant at all points of the container.

Properties/Assumptions of Ideal gas

- The molecules of a gas are in a state of continuous random motion. They move with all possible velocities in all possible directions. They obey Newton's law of motion.
- Mean momentum = 0; Mean velocity $\langle \vec{v} \rangle = 0$; $\langle v^2 \rangle \neq 0$ (Non zero); $\langle v^3 \rangle = \langle v^5 \rangle = 0$
- The average distance travelled by a molecule between two successive collisions is called as mean free path (λ_m) of the molecule.
- The time during which a collision takes place is negligible as compared to time taken by the molecule to cover the mean free path. At NTP ratio of time of collision to time of motion to cover mean free path is $10^{-8} : 1$.
- When a gas is taken into a vessel it is uniformly distributed in entire volume of vessel such that its mass density, molecular density, motion of molecules etc. all are identical for all direction, therefore root mean velocity

$$\overline{v_x^2} = \overline{v_y^2} = \overline{v_z^2} \text{ are equal. Pressure exerted by the gas in all direction } P_x = P_y = P_z = P \rightarrow \text{equal}$$

- All those assumptions can be justified, if number of gas molecules are taken very large i.e., 10^{23} molecules/cm³.

Different speeds of gas molecules

Average velocity

Because molecules are in random motion in all possible direction with all possible velocity. Therefore, the average velocity of the gas molecules in container is zero. $\langle \vec{v} \rangle = \frac{\vec{v}_1 + \vec{v}_2 + \dots + \vec{v}_N}{N} = 0$

$$\text{RMS speed of molecules } v_{rms} = \sqrt{\frac{3P}{\rho}} = \sqrt{\frac{3RT}{M_w}} = \sqrt{\frac{3kT}{m}} = 1.73 \sqrt{\frac{kT}{m}}$$

Average speed or Mean speed of molecules :

By maxwell's velocity distribution law v_M or $\langle |\vec{v}| \rangle \equiv v_{mean}$

$$\langle |\vec{v}| \rangle = v_{mean} = \frac{|\vec{v}_1| + |\vec{v}_2| + \dots + |\vec{v}_N|}{N} = \sqrt{\frac{8P}{\pi\rho}} = \sqrt{\frac{8RT}{\pi M_w}} = \sqrt{\frac{8kT}{\pi m}} = 1.59 \sqrt{\frac{kT}{m}}$$

Most probable speed of molecules (v_{mp})

At a given temperature, the speed to which maximum number of molecules belongs is called as most probable

$$\text{speed } (v_{mp}) \quad v_{mp} = \sqrt{\frac{2P}{\rho}} = \sqrt{\frac{2RT}{M_w}} = \sqrt{\frac{2kT}{m}} = 1.41 \sqrt{\frac{kT}{m}}$$

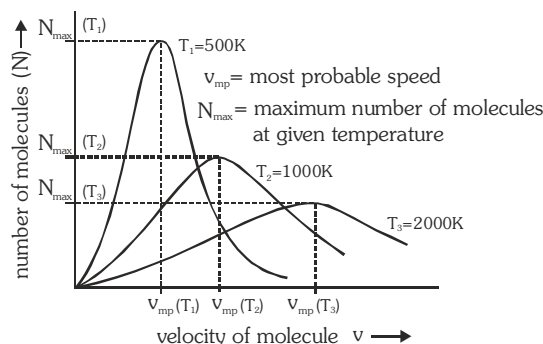
$$\text{Velocity of sound in gas medium } (v_s) \quad v_{\text{sound}} = \sqrt{\frac{\gamma P}{\rho}} = \sqrt{\frac{\gamma RT}{M_w}} = \sqrt{\frac{\gamma kT}{m}}$$

• At any temperature $v_{\text{rms}} > v_{\text{Mean}} > v_{\text{MP}} > v_{\text{sound}}$ (always)

• For a gas at any temperature (T) $\frac{v_{\text{rms}}}{v_{\text{sound}}} = \sqrt{\frac{3}{\gamma}}, \quad \frac{v_{\text{rms}}}{v_{\text{MP}}} = \sqrt{\frac{3}{2}}$

• A temperature is not possible at which above order can be changed

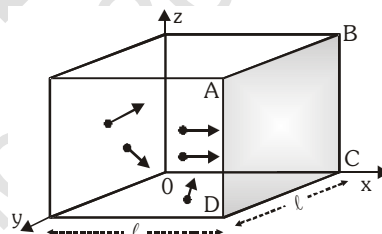
$$v_{\text{rms}} \neq v_{\text{Mean}} \neq v_{\text{MP}} \neq v_{\text{sound}} \quad (\text{always})$$

Maxwell's law of distribution of velocities**Expression for Pressure of An Ideal gas**

Consider an ideal gas enclosed in a cubical vessel of length ℓ .

Suppose there are 'N' molecules in a gas which are moving with

velocities $\vec{v}_1, \vec{v}_2, \dots, \vec{v}_N$.



If we consider any single molecule then its instantaneous velocity \vec{v} can be expressed as $\vec{v} = v_x \hat{i} + v_y \hat{j} + v_z \hat{k}$

Due to random motion of the molecule $v_x = v_y = v_z$; $|\vec{v}| = v_x \sqrt{3} = v_y \sqrt{3} = v_z \sqrt{3} = \sqrt{v_x^2 + v_y^2 + v_z^2}$

Suppose a molecule of mass m is moving with a velocity v_x towards the face ABCD. It strikes the face of the cubical vessel and returns back to strike the opposite face.

Change in momentum of the molecule per collision $\Delta p = -mv_x - mv_x = -2mv_x$

Momentum transferred to the wall of the vessel per molecule per collision $\Delta p = 2mv_x$

The distance travelled by the molecule in going to face ABCD and coming back is 2ℓ .

So, the time between two successive collision is $\Delta t = \frac{2\ell}{v_x}$

Number of collision per sec per molecule is $f_c = \frac{v_x}{2\ell} = \frac{\text{molecule velocity}}{\text{mean free path}}$; $f_c = \frac{v_{\text{rms}}}{\lambda_m}$ or $f_c = \frac{v_m}{\lambda_m}$

Hence momentum transferred to the wall per second by the molecule is equal to force applied on the wall

therefore, force $F = (2mv_x) \frac{v_x}{2\ell} = \frac{mv_x^2}{\ell} = \frac{mv^2}{3\ell}$ [As $v = \sqrt{3}v_x$]

$$\text{Pressure exerted by gas molecule } P = \frac{F}{A} = \frac{1}{3} \frac{mv^2}{\ell \times A} \Rightarrow P = \frac{1}{3} \frac{mv^2}{V} \quad [\because A \times \ell = V]$$

$$\text{pressure exerted by gas } P = \sum \frac{1}{3} \frac{mv^2}{V} = \sum \frac{1}{3} \frac{mv^2}{V} \times \frac{N}{N} = \frac{1}{3} \frac{mN}{V} \frac{\sum v^2}{N} = \frac{1}{3} \frac{mN}{V} v_{\text{rms}}^2$$

$$v_{\text{rms}}^2 = \frac{3PV}{M} = \frac{3\mu RT}{\mu M_w} \Rightarrow v_{\text{rms}} = \sqrt{\frac{3RT}{M_w}}, P = \frac{1}{3} \frac{M}{V} v_{\text{rms}}^2 = \frac{1}{3} \rho v_{\text{rms}}^2$$

- Average number of molecules for each wall = $\frac{N}{6}$.

$$\text{No. of molecules along each axis} = \frac{N}{3} \quad (N_x = N_y = N_z)$$

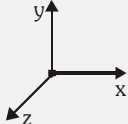
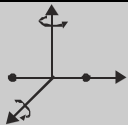
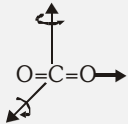

- $\bar{v}_x^2 = \bar{v}_y^2 = \bar{v}_z^2 = \frac{v_{\text{rms}}^2}{3}$ Root mean square velocity along any axis for gas molecule is $(v_{\text{rms}})_x = (v_{\text{rms}})_y = (v_{\text{rms}})_z = \frac{v_{\text{rms}}}{\sqrt{3}}$

All gas laws and gas equation can be obtained by expression of pressure of gas (except Joule's law)

Degree of freedom (f)

- The number of independent ways in which a molecule or an atom can exhibit motion or have energy is called its degrees of freedom.
- The number of independent coordinates required to specify the dynamic state of a system is called its degrees of freedom.
- The degrees of freedom are of three types :
 - Translational Degree of freedom** : There are maximum three degree of freedom corresponding to translational motion.
 - Rotational Degree of freedom** : The number of degrees of freedom in this case depends on the structure of the molecule.
 - Vibrational Degree of freedom** : It is exhibited at high temperature.

Degrees of freedom for different gases according to atomicity of gas at low temperature

Atomicity of gas	Translational	Rotational	Total	
Monoatomic Ex. Ar, Ne, Ideal gas etc	3	0	3	
Diatomic Ex. O ₂ , Cl ₂ , N ₂ etc.	3	2	5	
Triatomic (linear) Ex. CO ₂ , C ₂ H ₂	3	2	5	
Triatomic (Non-linear) or Polyatomic Ex. H ₂ O, NH ₃ , CH ₄	3	3	6	

At high temperature a diatomic molecule has 7 degrees of freedom. (3 translational, 2 rotational and 2 vibrational)

Maxwell's law of equipartition of energy

The total kinetic energy of a gas molecules is equally distributed among its all degree of freedom and the energy associated with each degree of freedom at absolute temperature T is $\frac{1}{2}kT$

For one molecule of gas

Energy related with each degree of freedom = $\frac{1}{2}kT$

Energy related with all degree of freedom = $\frac{f}{2}kT$ $\because \bar{v}_x^2 = \bar{v}_y^2 = \bar{v}_z^2 = \frac{v_{rms}^2}{3} \Rightarrow \frac{1}{2}mv_{rms}^2 = \frac{3}{2}kT$

So energy related with one degree of freedom = $\frac{1}{2}m \frac{v_{rms}^2}{3} = \frac{3}{2} \frac{kT}{3} = \frac{1}{2}kT$

Different K.E. of gas (Internal Energy)

- Translational kinetic energy (E_T)** $E_T = \frac{1}{2}Mv_{rms}^2 = \frac{3}{2}PV$

Kinetic energy of volume V is = $\frac{1}{2}Mv_{rms}^2$ [Note : Total internal energy of ideal gas is total kinetic energy]

- Energy per unit volume or energy density (E_v)** $E_v = \frac{\text{Total energy}}{\text{Volume}} = \frac{E}{V}$; $E_v = \frac{1}{2} \left[\frac{M}{V} \right] v_{rms}^2 = \frac{1}{2} \rho v_{rms}^2$

$$\therefore P = \frac{2}{3} \left[\frac{1}{2} \rho v_{rms}^2 \right] \therefore E_v = \frac{3}{2}P$$

- Molar K.E. or Mean Molar K.E. (E)** $E = \frac{1}{2}M_w v_{rms}^2$ for N_0 molecules or M_w (gram) ;

$$E = \frac{3}{2}RT = \frac{3}{2}N_0kT$$

- Molecular kinetic energy or mean molecular K.E. (\bar{E})** $E = \frac{1}{2}M_w v_{rms}^2$, $\bar{E} = \frac{E}{N_0} = \frac{3}{2} \frac{RT}{N_0} = \frac{3}{2}kT$

- Kinetic energy of 1 g mass (E_m)** $= \frac{3}{2} \frac{RT}{M_w} = \frac{3}{2} \frac{N_0 \times kT}{N_0 m} = \frac{3}{2} \frac{kT}{m}$

SPECIFIC HEAT CAPACITY**Monoatomic Gases**

The molecule of a monoatomic gas has only three translational degrees of freedom. Thus, the average energy of a molecule at temperature T is $(3/2)kT$.

The total internal energy of a mole of such a gas is

$$U = \frac{3}{2}kT \times N_A = \frac{3}{2}RT$$

The molar specific heat at constant volume, C_v , is

$$C_v (\text{monoatomic gas}) = \frac{dU}{dT} = \frac{3}{2}R$$

Diatomic Gases

As explained earlier, a diatomic molecule treated as a rigid rotator like a dumbbell has 5 degrees of freedom : 3 translational and 2 rotational. Using the law of equipartition of energy, the total internal energy of a mole of such a gas is

$$U = \frac{5}{2}kT \times N_A = \frac{5}{2}RT$$

The molar specific heats are then given by

$$C_v (\text{rigid diatomic}) = \frac{5}{2}R$$

MEAN FREE PATH

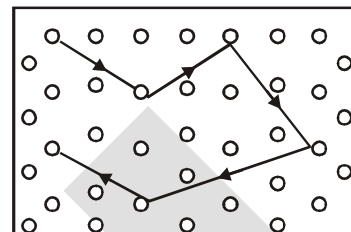
A molecule in its path undergoes a number of collisions so the path traversed by it is not a straight line but somewhat zigzag. Between two successive collisions a molecule travels in a straight line with uniform velocity. As motion is random thus the distance travelled by molecule between two successive collisions is not always equal.

The average distance travelled by a molecule between two successive collisions is called as mean free path (λ_m) of the molecule.

$$\lambda_m = \frac{1}{\sqrt{2}\pi d^2 n}$$

Here $n = \frac{N}{V}$ = No. of molecules per unit volume

d = diameter of a molecule



GOLDEN KEY POINTS

- At a given temperature graph drawn between molecular velocity and number of molecules is known as velocity distribution curve.
- The velocities of molecules of a gas are in between zero and infinity ($0 - \infty$)
- With the increase in the temperature, the most probable velocity and maximum molecule velocity both increases.
- The number of molecules within certain velocity range is constant although the velocity of molecule changes continuously at particular temperature.
- The area enclosed between the ($N - v$) curve and the velocity axis represents the total number of molecules.

On the basis of velocity distribution Maxwell established the law of equipartition of energy for gases at any temperature.

- Except 0 K, at any temperature T , $E > E_m > \bar{E}$
- At a common temperature, for all ideal gas

E and \bar{E} are same while E_m is different and depends upon nature of gas (M_w or m)

- For thermal equilibrium of gases, temperature of each gas is same and this temperature called as temperature of mixture (T_m) which can be find out on basis of conservation of energy (All gases are of same atomicity).

$$T_m = \frac{\sum NT}{\sum N} = \frac{N_1 T_1 + N_2 T_2 + \dots + N_n T_n}{N_1 + N_2 + \dots + N_n}$$

- 1 mole gas :** Mean translational kinetic energy = $\frac{3}{2}RT$; Total kinetic energy = $\frac{f}{2}RT$

1 molecule of gases : Mean translational kinetic energy = $\frac{3}{2}kT$; Total kinetic energy = $\frac{f}{2}kT$;

$f \rightarrow$ Degree of freedom

- Specific heats of all substances approach zero as $T \rightarrow 0$. This is related to the fact that degrees of freedom get frozen and ineffective at low temperatures.

Illustrations

Illustration 48.

The velocities of ten particles in ms^{-1} are 0, 2, 3, 4, 4, 4, 5, 5, 6, 9. Calculate

- (i) average speed and (ii) rms speed (iii) most probable speed.

Solution

$$(i) \quad \text{average speed, } v_{av} = \frac{0+2+3+4+4+4+5+5+6+9}{10}$$

$$= \frac{42}{10} = 4.2 \text{ ms}^{-1}$$

$$(ii) \quad \text{rms speed, } v_{rms} = \left[\frac{(0)^2 + (2)^2 + (3)^2 + (4)^2 + (4)^2 + (4)^2 + (5)^2 + (5)^2 + (6)^2 + (9)^2}{10} \right]^{1/2}$$

$$= \left[\frac{228}{10} \right]^{1/2} = 4.77 \text{ ms}^{-1}$$

$$(iii) \quad \text{most probable speed } v_{mp} = 4 \text{ m/s}$$

Illustration 49.

At what temperature root mean square velocity of hydrogen becomes double of its value at S.T.P., pressure remaining constant ?

Solution

Let v_1 be the r.m.s. velocity at S.T.P. and v_2 be the r.m.s. velocity at unknown temperature T_2 .

$$\therefore \frac{v_1^2}{v_2^2} = \frac{T_1}{T_2}$$

$$\text{or } T_2 = T_1 \left[\frac{v_2}{v_1} \right]^2 = 273 \times (2)^2 = 273 \times 4 = 1092 \text{ K} = (1092 - 273) = 819^\circ\text{C}$$

Illustration 50.

Calculate rms velocity of oxygen molecule at 27°C

Solution

Temperature, $T = 27^\circ\text{C}$

$$\Rightarrow 273 + 27 = 300 \text{ K,}$$

Molecular weight of oxygen = $32 \times 10^{-3} \text{ kg}$ and $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$

$$\text{rms velocity is } v_{rms} = \sqrt{\frac{3RT}{M}} = \sqrt{\frac{3 \times 8.31 \times 300}{32 \times 10^{-3}}} = 483.5 \text{ ms}^{-1}$$

Illustration 51.

Calculate the kinetic energy of a gram molecule of argon at 127°C .

Solution

Temperature, $T = 127^\circ\text{C} = 273 + 127 = 400 \text{ K, } R = 8.31 \text{ J/mol K}$

$$\text{K.E. per gram molecule of argon} = \frac{3}{2} R T = \frac{3}{2} \times 8.31 \times 400 = 4986 \text{ J}$$

Illustration 52.

The mass of a hydrogen molecule is 3.32×10^{-27} kg. If 10^{23} molecules are colliding per second on a stationary wall of area 2 cm^2 at an angle of 45° to the normal to the wall and reflected elastically with a speed 10^3 m/s. Find the pressure exerted on the wall. (in N/m^2)

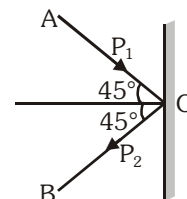
Solution

As the impact is elastic $\therefore |\vec{p}_1| = |\vec{p}_2| = p = mv = 3.32 \times 10^{-24} \text{ kg m/s}$

The change in momentum along the normal $\Delta p = |\vec{p}_2 - \vec{p}_1| = 2p \cos 45^\circ = \sqrt{2}p$

If f is the collision frequency then force applied on the wall $F = \frac{\Delta p}{\Delta t} = \Delta p \times f = \sqrt{2}pf$

$$\therefore \text{Pressure } P = \frac{F}{A} = \frac{\sqrt{2}pf}{A} = \frac{\sqrt{2} \times 3.32 \times 10^{-24} \times 10^{23}}{2 \times 10^{-4}} = 2.347 \times 10^3 \text{ N/m}^2$$



BEGINNER'S BOX-6

- Given : Avogadro number $N = 6.02 \times 10^{23}$ molecules/mole and Boltzmann's constant $k = 1.38 \times 10^{-23} \text{ J/(molecule-K)}$.
Calculate :
 (a) The average kinetic energy of translation of an oxygen molecule at 27°C
 (b) The total kinetic energy of an oxygen molecule at 27°C
 (c) The total kinetic energy of 1 mole of oxygen gas at 27°C .
- The temperature of gas is -73°C . To what temperature should it be heated so that
 (a) average kinetic energy of the molecule is doubled ?
 (b) the root mean square velocity of the molecules is doubled ?
- For a gas $\frac{R}{C_p} = 0.4$. For this gas calculate the following -
 (a) atomicity and degree of freedom
 (b) value of C_v and γ
 (c) mean gram - molecular kinetic energy at 300 K temperature
- If three molecules are having speeds v_1, v_2 and v_3 respectively, then what will be their average speed and root mean square speed ?
- Four molecules of a gas are having speeds of $1, 4, 8$ and 16 ms^{-1} . Find the root mean square velocity of the gas molecules ?
- A flask contains argon and chlorine in the ratio of $2 : 1$ by mass. The temperature of the mixture is 27°C . Obtain the ratio of
 (a) Average translational kinetic energy per molecule, and
 (b) root mean square speed (v_{rms}) of the molecules of the two gases.
 Atomic mass of argon = 39.9 u :
 Molecular mass of chlorine = 70.9 u .

THERMODYNAMICS

Branch of physics which deals with the inter-conversion between heat energy and any other form of energy is known as thermodynamics. In this branch of physics we deal with the processes involving heat, work and internal energy. In this branch of science the conversion of heat into mechanical work and vice versa is studied.

- **Thermodynamic System**

The system which can be represented in terms of pressure (P), volume (V) and temperature (T), is known as thermodynamic system. A specified portion of matter consisting of one or more substances on which the effects of thermodynamic variables such as temperature, volume and pressure are to be studied, is called a system.

e.g. A gas enclosed in a cylinder fitted with a piston is a system.

- **Surrounding**

Anything outside the system, which exchanges energy with the system and which tends to change the properties of the system is called its surrounding.

- **Heterogeneous System**

A system which is not uniform throughout is said to be heterogeneous. e.g. A system consisting of two or more immiscible liquids.

- **Homogeneous System**

A system is said to be homogeneous if it is completely uniform throughout. e.g. Pure solid or liquid.

- **Isolated System**

A system in which no exchange of matter and energy with the surrounding takes place, is said to be an isolated system.

- **Universe**

The system and its surrounding are together known as the universe.

- **Thermodynamic variables of the system**

(i) Composition (μ) (ii) Temperature (T) (iii) Volume (V) (iv) Pressure (P) (v) Mass

- **Thermodynamic state**

The state of a system can be described completely by composition, temperature, volume and pressure.

If a system is homogeneous and has definite mass and composition, then the state of the system can be described by the remaining three variables namely temperature, pressure and volume. These variables are inter related by equation $PV = \mu RT$. The thermodynamic state of the system is its condition as identified by two independent thermodynamic variables (P, V or P, T or V, T).

- **Internal Energy**

Internal energy of a system is the energy possessed by the system due to molecular motion and molecular configuration. The energy due to molecular motion is called internal kinetic energy (U_k) and that due to molecular configuration is called internal potential energy (U_p). $dU = dU_k + dU_p$

If there are no intermolecular forces, then $dU_p = 0$ and $dU = dU_k = m c_v dT$ [ideal gas]

c_v = Specific heat at constant volume and dT = Infinitesimal change in temperature

m = Mass of system

M = Molecular weight

Molar heat capacity $C_v = M c_v$ For μ -moles of ideal gas $dU = \mu C_v dT = \frac{m}{M} C_v dT$

Internal energy in the absence of inter-molecular forces is simply the function of temperature and state only, it is independent of path followed. $\Delta U = U_f - U_i$

U_i = Internal energies in initial state and U_f = Internal energies in final state

Thermodynamic Processes

In the thermodynamic process pressure, volume, temperature and entropy of the system change with time. Thermodynamic process is said to take place if change occurs in the state of a thermodynamic system.

Sign convention used for the study of thermodynamic processes

Heat gained by a system	Positive
Heat lost by a system	Negative
Work done by a system	Positive
Work done on the system	Negative
Increase in the internal energy of system	Positive
Decrease in the internal energy of system	Negative

Indicator Diagram or P-V Diagram

In the equation of state of a gas $PV = \mu RT$ two thermodynamic variables are sufficient to describe the behavior of a thermodynamic system.

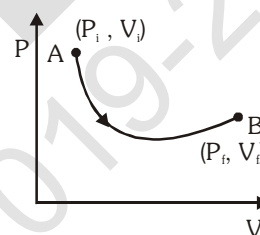
If any two of the three variables P , V and T are known then the third can be calculated.

P - V diagram is a graph between the volume V and the pressure P of the system.

The volume is plotted against X-axis while the pressure is plotted against Y-axis.

The point A represents the initial stage of the system. Initial pressure of the system is P_i and initial volume of the system V_i .

The point B represents the final state of the system. P_f and V_f are the final pressure and final volume respectively of the system. The points between A and B represent the intermediate states of the system. With the help of the indicator diagram we calculate the amount of work done by the gas or on the gas during expansion or compression.



Cyclic process

Cyclic process is a thermodynamic process in which the system returns to its initial stage after undergoing a series of changes.

Non-cyclic process

Non-cyclic process is a process in which the system does not return to its initial stage.

Quasi-static or equilibrium process

Quasi-static is a thermodynamic process which proceeds extremely slowly such that at every instant of time, the temperature and pressure are the same in all parts of the system.

Reversible and Irreversible processes

A reversible process is one in which the changes in heat and work of direct process from initial to a final state are exactly retraced in opposite sense in the reverse process and the system and surroundings are left in their initial states. The reversibility is an ideal concept and can not be realized in practice.

The process which is not reversible is the irreversible process. In nature the processes are irreversible.

Work done by thermodynamic system

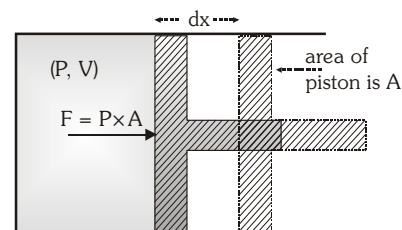
One of the simple example of a thermodynamic system is a gas in a cylinder with a movable piston.

If the gas expands against the piston

Gas exerts a force on the piston and displace it through a distance and does work on the piston.

If the piston compresses the gas

When piston moved inward, work is done on the gas.



• **The work associated with volume changes**

If pressure of gas on the piston = P .

Then the force on the piston due to gas is $F = PA$

When the piston is pushed outward an infinitesimal distance dx ,

the work done by the gas is $dW = F \times dx = PA \, dx$

The change in volume of the gas is $dV = A \, dx$, $\therefore dW = P \, dV$

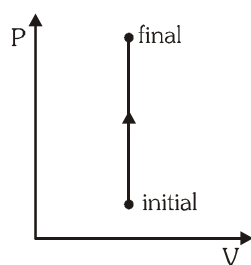
For a finite change in volume from V_i to V_f , this equation is then integrated between V_i to V_f to find the

$$\text{net work done } W = \int_{V_i}^{V_f} dW = \int_{V_i}^{V_f} P \, dV$$

Hence the work done by a gas is equal to the area under P - V graph.

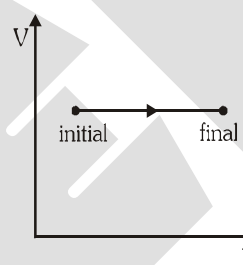
Following different cases are possible.

(i) Volume is constant



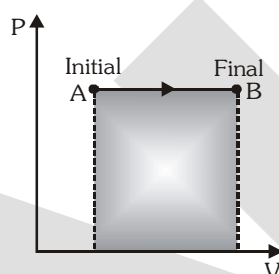
$V = \text{constant}$ and

or



$W = 0$

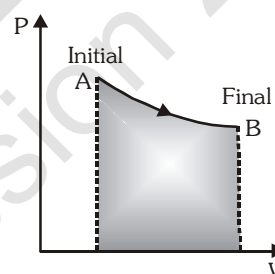
(ii) Volume is increasing



V is increasing

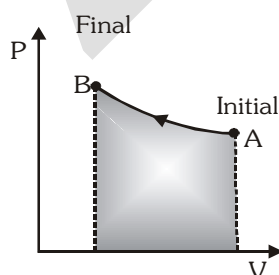
$W_{AB} > 0$

or



$W_{AB} = \text{Shaded area}$

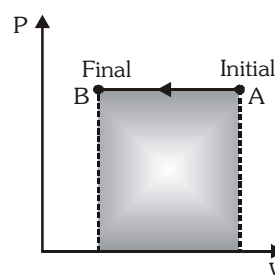
(iii) Volume is decreasing



V is decreasing

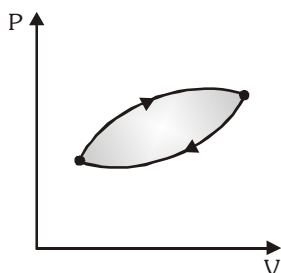
$W_{AB} < 0$

or

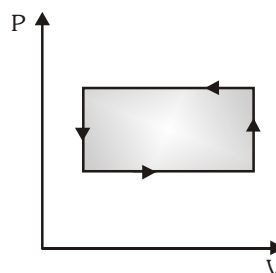


$W_{AB} = - \text{Shaded area}$

(iv) Cyclic process

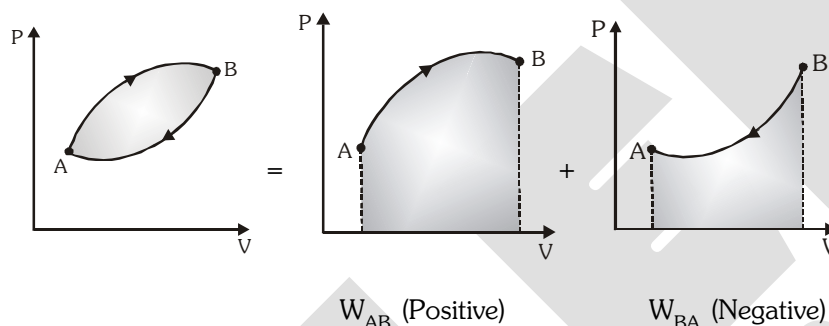


$$W_{\text{clockwise cycle}} = + \text{Shaded area}$$



$$W_{\text{anticlockwise cycle}} = - \text{Shaded area}$$

Work Done in Clockwise Cycle



$$W_{\text{cyclic}} = W_{AB} (\text{positive}) + W_{BA} (\text{negative}) =$$



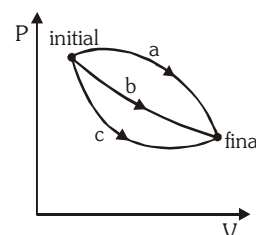
First law of thermodynamics

Let a gas in a cylinder with a movable piston changes from an initial equilibrium state to a final equilibrium state.

System change its state through path 'a' :

The heat absorbed by the system in this process = δQ_a

The work done by the system = δW_a



Again for path 'b' :

Heat absorbed by the system = δQ_b , Work done by the system = δW_b .

It is experimental fact that the $\delta Q_a - \delta W_a = \delta Q_b - \delta W_b$

Both δQ and δW depend on the thermodynamic path taken between two equilibrium states, but difference ($\delta Q - \delta W$) does not depends on path in between two definite states of the system.

So, there is a function (internal energy) of the thermodynamic coordinates (P , V and T) whose final value (U_f) minus its initial value (U_i) equals the change $\delta Q - \delta W$ in the process.

$dU = \delta Q - \delta W$. This is the first law of thermodynamics.

Heat supplied to the system and work done by the system are path dependent so they are denoted by δQ and δW respectively. Change in internal energy $\Delta U = U_f - U_i$ does not depends on path it depends only on initial and final positions of the system. So, it is denoted by dU (or ΔU)

First Law of Thermodynamics

If some quantity of heat is supplied to a system capable of doing external work, then the quantity of heat absorbed by the system is equal to the sum of the increase in the internal energy of the system and the external work done by the system.

$$\boxed{\delta Q = dU + \delta W} \quad \text{or} \quad \boxed{Q = W + \Delta U}$$

- This law is applicable to every process in nature
- The first law of thermodynamics introduces the concept of internal energy.
- The first law of thermodynamics is based on the law of conservation of energy.
- δQ , dU and δW must be expressed in the same units (either in units of work or in units of heat).
- This law is applicable to all the three phases of matter, i.e., solid, liquid and gas.
- dU is a characteristic of the state of a system, it may be any type of internal energy—translational kinetic energy, vibrational, rotational kinetic energy, binding energy etc.

Application of First Law of Thermodynamics

- **Melting Process :**

When a substance melts, the change in volume (dV) is very small and can, therefore, be neglected. The temperature of a substance remains unchanged during melting process.

Let us consider the melting of a mass m of the solid. Let L be the latent heat of fusion i.e., the heat required L to change a unit mass of a solid to liquid phase at constant temperature.

Heat absorbed during melting process, $Q = mL$

By the first law of thermodynamics, $Q = \Delta U + W = \Delta U + P\Delta V$, where ΔV is volume change during phase change.

- **Boiling Process :**

When a liquid is heated, it changes into vapour at constant temperature (called boiling point) and pressure. When water is heated at normal atmospheric pressure, it boils at 100°C . The temperature remains unchanged during the boiling process. Let us consider the vaporisation of liquid of mass m . Let V_ℓ and V_v be the volumes of the liquid and vapours respectively.

The work done in expanding at constant temperature and pressure P , $\delta W = P\Delta V = P(V_v - V_\ell)$

Let the latent heat of vaporisation = L \therefore Heat absorbed during boiling process, $\delta Q = mL$

Let U_ℓ and U_v be the internal energies of the liquid and vapours respectively then $\Delta U = U_v - U_\ell$

According to the first law of thermodynamics, $\delta Q = \Delta U + \delta W \therefore mL = (U_v - U_\ell) + P(V_v - V_\ell)$

Illustrations**Illustration 53.**

The pressure of one mole monoatomic gas increases linearly from $4 \times 10^5 \text{ Nm}^{-2}$ to $8 \times 10^5 \text{ Nm}^{-2}$ when its volume increases from 0.2 m^3 to 0.5 m^3 . Calculate.

- | | |
|--------------------------------|---------------------------------------|
| (i) Work done by the gas, | (ii) Increase in the internal energy, |
| (iii) Amount of heat supplied, | (iv) Molar heat capacity of the gas |

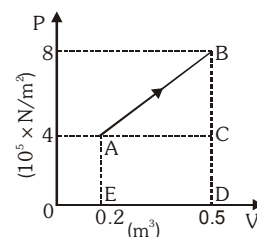
$$R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$$

Solution

$$P_1 = 4 \times 10^5 \text{ Nm}^{-2} \quad P_2 = 8 \times 10^5 \text{ Nm}^{-2}, \quad V_1 = 0.2 \text{ m}^3, \quad V_2 = 0.5 \text{ m}^3$$

- (i) Work done by the gas = Area under P - V graph (Area ABCDEA)

$$\begin{aligned}
 &= \frac{1}{2} (AE + BD) \times AC = \frac{1}{2} (4 \times 10^5 + 8 \times 10^5) \times (0.5 - 0.2) \\
 &= \frac{1}{2} \times 12 \times 10^5 \times 0.3 = 1.8 \times 10^5 \text{ J}
 \end{aligned}$$



$$(ii) \text{ Increase in internal energy } \Delta U = C_V (T_2 - T_1) = \frac{C_V}{R} R(T_2 - T_1) = \frac{C_V}{R} (P_2 V_2 - P_1 V_1)$$

$$\text{For monoatomic gas } C_V = \frac{3}{2}R$$

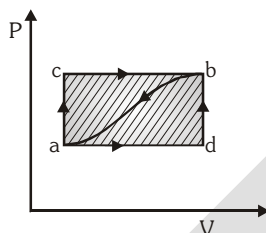
$$\therefore \Delta U = \frac{3}{2} [(8 \times 10^5 \times 0.5) - (4 \times 10^5 \times 0.2)] = \frac{3}{2} [4 \times 10^5 - 0.8 \times 10^5] = 4.8 \times 10^5 \text{ J}$$

$$(iii) Q = \Delta U + W = 4.8 \times 10^5 + 1.8 \times 10^5 = 6.6 \times 10^5 \text{ J}$$

$$(iv) C = \frac{Q}{\mu \Delta T} = \frac{QR}{\mu R \Delta T} = \frac{QR}{(P_2 V_2 - P_1 V_1)} = \frac{6.6 \times 10^5 \times 8.31}{1 \times 3.2 \times 10^5} = 17.14 \text{ J/mol K}$$

Illustration 54.

As shown in figure when a system is taken from state a to state b, along the path



a \rightarrow c \rightarrow b, 60 J of heat flow into the system, and 30 J of work is done :

- How much heat flows into the system along the path a \rightarrow d \rightarrow b if the work is 10 J.
- When the system is returned from b to a along the curved path, the work done by the system is -20 J. Does the system absorb or liberate heat, and how much ?
- If, $U_a = 0$ and $U_d = 22$ J, find the heat absorbed in the process a \rightarrow d and d \rightarrow b.

Solution

$$\text{For the path acb} \quad \Delta U = Q - W = 60 - 30 = 30 \text{ J or } U_b - U_a = 30 \text{ J}$$

$$(i) \text{ Along the path adb} \quad Q = \Delta U + W = 30 + 10 = 40 \text{ J}$$

$$(ii) \text{ Along the curved path ba} \quad Q = (U_a - U_b) + W = (-30) + (-20) = -50 \text{ J, heat liberates from system}$$

$$(iii) Q_{ad} = U_d - U_a + W_{ad}$$

$$\text{but } W_{ad} = W_{adb} - W_{db} = 10 - 0 = 10 \quad \text{Hence } Q_{ad} = 22 - 0 + 10 = 32 \text{ J}$$

$$\text{and } Q_{db} = U_b - U_d + W_{db} = 30 - 22 + 0 = 8 \text{ J}$$

Illustration 55.

1 kg of water at 373 K is converted into steam at same temperature. Volume of 1 cm³ of water becomes 1671 cm³ on boiling. What is the change in the internal energy of the system if the latent heat of vaporisation of water is 5.4×10^5 cal kg⁻¹?

Solution

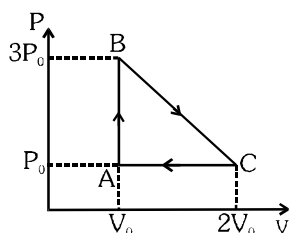
$$\text{Volume of 1 kg of water} = 1000 \text{ cm}^3 = 10^{-3} \text{ m}^3, \text{ Volume of 1 kg of steam} = 10^3 \times 1671 \text{ cm}^3 = 1.671 \text{ m}^3$$

$$\text{Change in volume, } \Delta V = (1.671 - 10^{-3}) \text{ m}^3 = 1.670 \text{ m}^3, \text{ Pressure, } P = 1 \text{ atm.} = 1.01 \times 10^5 \text{ N m}^{-2}$$

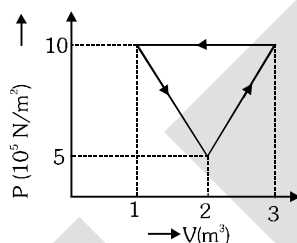
$$\text{In expansion work done, } W = P \Delta V = 1.01 \times 10^5 \times 1.67 \text{ J} = \frac{1.686 \times 15^5}{4.2} \text{ cal} = 4.015 \times 10^4 \text{ cal}$$

$$\text{But } \Delta U = Q - W \text{ (first law of thermodynamics) or } \Delta U = (5.4 \times 10^5 - 0.4015 \times 10^5) \text{ cal} = 4.9985 \times 10^5 \text{ cal}$$

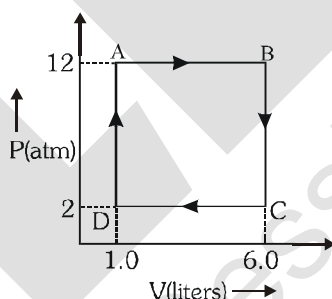
1. One mole of an ideal monoatomic gas is taken round the cyclic process ABCA as shown in fig. calculate.



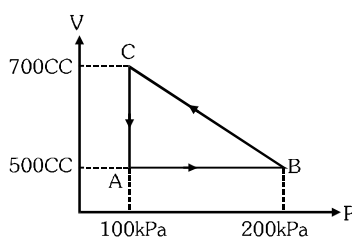
- The work done by the gas.
 - The heat rejected by the gas in the path CA and heat absorbed in the path AB.
 - The net heat absorbed by the gas in the path BC.
2. Calculate the heat supplied in figure shown for complete cycle.



3. Given figure shows a P-V graph of the thermodynamic behaviour of an ideal gas. Find out work in the processes A → B, B → C, C → D and D → A from this graph.



4. At 8 atm pressure a monoatomic gas expands from 2 litre to 5 litre then calculate the following
- Work done by the gas
 - Increase in internal energy
 - Amount of heat supplied
5. A gas is taken through a cyclic process ABCA as shown in figure. If 2.4 cal of heat is given in the process. Calculate the value of J.



6. DIFFERENT PROCESSES

6.1 Isometric or Isochoric Process

Isochoric process is a thermodynamic process that takes place at constant volume of the system, but pressure and temperature varies for change in state of the system.

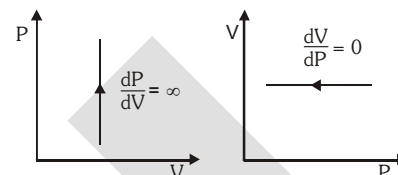
Equation of state $P/T = \text{constant}$ (P and T are variable, V is constant)

Work done In this process volume remains constant so $dV = 0 \Rightarrow W = \int_{V_i}^{V_f} PdV = 0$

Form of first Law $Q = \Delta U$

It means whole of the heat supplied is utilized for change in internal energy of the system. $Q = \Delta U = \mu C_v \Delta T$

Slope of the P-V curve $\frac{dP}{dV} = \infty$



Specific heat or molar heat capacity at constant volume (C_v)

The quantity of heat required to raise the temperature of 1 gram mole of gas through 1°C at constant volume is equal to the specific heat at constant volume.

- A gas enclosed in a cylinder having rigid walls and a fixed piston. When heat is added to the gas, there would be no change in the volume of the gas.
- When a substance melts, the change in volume is negligibly small. So, this may be regarded as a nearly isochoric process.
- Heating process in pressure cooker is an example of isometric process.

6.2 Isobaric Process

Isobaric process is a thermodynamic process that takes place at constant pressure, but volume and temperature varies for change in state of the system.

- **Equation of state** $V/T = \text{Constant}$ or $V \propto T$
- **Work done** In this process pressure remains constant $\therefore dP = 0$

$$\text{Work done } W = \int_{V_i}^{V_f} PdV = P(V_f - V_i)$$

\Rightarrow Fraction of Heat given which -

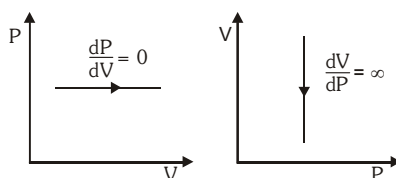
(i) is converted into internal energy $\frac{\Delta U}{\Delta Q} = \frac{1}{\gamma}$ (ii) does work against external pressure $\frac{\Delta W}{\Delta Q} = 1 - \frac{1}{\gamma}$

- **Form of first Law** $Q = \Delta U + P(V_f - V_i)$
 $\mu C_p (T_f - T_i) = \mu C_v (T_f - T_i) + P(V_f - V_i)$

It is clear that heat supplied to the system is utilized for :

(i) Increasing internal energy and (ii) Work done against the surrounding atmosphere.

- **Slope of the PV curve :** $\left(\frac{dP}{dV}\right)_{\text{isobaric}} = 0$



- **Specific heat at constant pressure (C_p)**

The quantity of heat required to raise the temperature of 1 gram mole gas through 1°C at constant pressure is equal to the specific heat.

Ex. Heating of water at atmospheric pressure. • Melting of solids and boiling of liquids at atmospheric pressure.

6.3 Isothermal Process

In this process pressure and volume of system change but temperature remains constant. In an isothermal process, the exchange of heat between the system and the surroundings is allowed. Isothermal process is carried out by either supplying heat to the substance or by extracting heat from it. A process has to be extremely slow to be isothermal.

- **Equation of state :** $PV = \text{constant}$ (μRT) (T is constant)

- **Work Done**

Consider μ moles of an ideal gas, enclosed in a cylinder, at absolute temperature T , fitted with a frictionless piston. Suppose that gas undergoes an isothermal expansion from the initial state (P_1, V_1) to the final state

(P_2, V_2). Work done : $W = \int_{V_1}^{V_2} PdV$ $\because PV = \mu RT$ then $P = \frac{\mu RT}{V}$

$$\therefore W = \int_{V_1}^{V_2} \frac{\mu RT}{V} dV = \mu RT \int_{V_1}^{V_2} \frac{dV}{V} = \mu RT [\log_e V]_{V_1}^{V_2} = \mu RT [\log_e V_2 - \log_e V_1] = \mu RT \log_e \left[\frac{V_2}{V_1} \right]$$

$$\Rightarrow W = 2.303 \mu RT \log_{10} \left[\frac{P_1}{P_2} \right] \quad [\because P_1 V_1 = P_2 V_2]$$

Form of First Law

There is no change in temperature and internal energy of the system depends on temperature only

$$\text{So } \Delta U = 0, Q = 2.303 \mu RT \log_{10} \left[\frac{V_2}{V_1} \right]$$

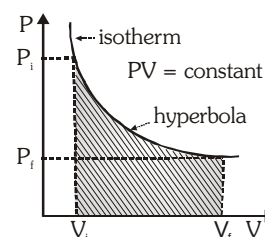
It is clear that Whole of the heat energy supplied to the system is utilized by the system in doing external work. There is no change in the internal energy of the system.

Slope of the isothermal curve

For an isothermal process, $PV = \text{constant}$

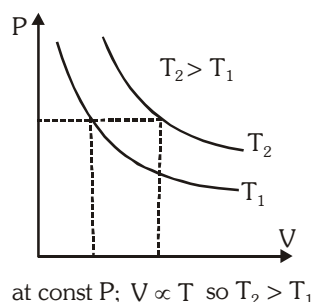
$$\text{Differentiating, } PdV + VdP = 0 \Rightarrow VdP = -PdV \Rightarrow \frac{dP}{dV} = -\frac{P}{V}$$

$$\text{Slope of isothermal curve, } \left[\frac{dP}{dV} \right]_{\text{isothermal}} = -\frac{P}{V}$$



For a given system :

- The product of the pressure and volume of a given mass of a perfect gas remains constant in an isothermal process.
- Boyle's law is obeyed in an isothermal process.
- A graph between pressure and volume of a given mass of a gas at constant temperature is known as isotherm or isothermal curve of the gas.
- Two isotherms for a given gas at two different temperatures T_1 and T_2 are shown in figure.



- The curves drawn for the same gas at different temperatures are mutually parallel and do not cut each other.
- If two isotherms intersect each other at a single point we get same value of P and V at intersection point. $PV = \mu RT_1$ for temperature T_1 and $PV = \mu RT_2$ for temperature T_2 .
It means $T_1 = T_2$ which is not possible.
- An ideal gas enclosed in a conducting cylinder fitted with a conducting piston.
- Let the gas be allowed to expand very-very slowly. This shall cause a very slow cooling of the gas, but heat will be conducted into the cylinder from the surrounding. Hence the temperature of the gas remains constant. If the gas is compressed very-very slowly, heat will be produced, but this heat will be conducted to the surroundings and the temperature of the gas shall remain constant.
- The temperature of a substance remains constant during melting. So, the melting process is an isothermal process.
- Boiling is an isothermal process, when a liquid boils, its temperature remains constant.
- If sudden changes are executed in a vessel of infinite conductivity then they will be isothermal.

6.4 Adiabatic Process

It is that thermodynamic process in which pressure, volume and temperature of the system do change but there is no exchange of heat between the system and the surroundings.

A sudden and quick process will be adiabatic since there is no sufficient time available for exchange of heat so process adiabatic.

Equation of state : $PV = \mu RT$

Equation for adiabatic process $PV^\gamma = \text{constant}$, $TV^{\gamma-1} = \text{constant}$, $T^\gamma P^{1-\gamma} = \text{constant}$

Work done

Let initial state of system is (P_1, V_1, T_1) and after adiabatic change final state of system is (P_2, V_2, T_2) then we can write $P_1 V_1^\gamma = P_2 V_2^\gamma = K$ (here K is const.)

$$\text{So } W = \int_{V_1}^{V_2} P dV = K \int_{V_1}^{V_2} V^{-\gamma} dV = K \left(\frac{V^{-\gamma+1}}{-\gamma+1} \right)_{V_1}^{V_2} = \frac{K}{(-\gamma+1)} [V_2^{-\gamma+1} - V_1^{-\gamma+1}] \quad (\because K = P_1 V_1^\gamma = P_2 V_2^\gamma)$$

$$\Rightarrow W = \frac{1}{(\gamma-1)} [P_1 V_1^\gamma V_1^{-\gamma} \cdot V_1 - P_2 V_2^\gamma V_2^{-\gamma} \cdot V_2] = \frac{1}{(\gamma-1)} [P_1 V_1 - P_2 V_2] \Rightarrow W = \frac{\mu R}{(\gamma-1)} (T_1 - T_2) \quad (\because PV = \mu RT)$$

Form of first law : $dU = -\delta W$

It means the work done by an ideal gas during adiabatic expansion (or compression) is on the cost of change in internal energy proportional to the fall (or rise) in the temperature of the gas.

If the gas expands adiabatically, work is done by the gas. So, W_{adia} is positive.

The gas cools during adiabatic expansion and $T_1 > T_2$.

If the gas is compressed adiabatically, work is done on the gas. So, W_{adia} is negative.

The gas heats up during adiabatic compression and $T_1 < T_2$.

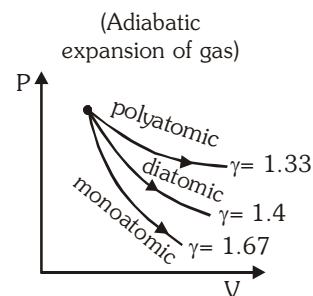
Slope of the adiabatic curve

For an adiabatic process, $PV^\gamma = \text{constant}$

Differentiating, $\gamma P V^{\gamma-1} dV + V^\gamma dP = 0$

$$\Rightarrow V^\gamma dP = -\gamma P V^{\gamma-1} dV \Rightarrow \frac{dP}{dV} = -\frac{\gamma P V^{\gamma-1}}{V^\gamma} = -\gamma \frac{P}{V} = \gamma \left(-\frac{P}{V} \right)$$

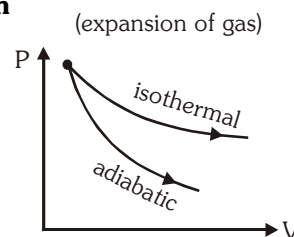
$$\text{Slope of adiabatic curve, } \left[\frac{dP}{dV} \right]_{\text{adiabatic}} = -\frac{\gamma P}{V}$$



Magnitude of slope of adiabatic curve is greater than the slope of isotherm

$$\left| \frac{dP}{dV} \right|_{\text{adia}} = \gamma \frac{P}{V} = \gamma \left| \frac{dP}{dV} \right|_{\text{iso}} \Rightarrow \frac{\text{slope of adiabatic changes}}{\text{slope of isothermal changes}} = \gamma$$

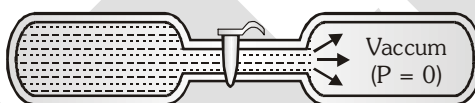
Since γ is always greater than one so an adiabatic curve is steeper than an isothermal curve

**Examples of adiabatic process**

- A gas enclosed in a thermally insulated cylinder fitted with a non-conducting piston. If the gas is compressed suddenly by moving the piston downwards, some heat is produced. This heat cannot escape from the cylinder. Consequently, there will be an increase in the temperature of the gas.
- If a gas is suddenly expanded by moving the piston outwards, there will be a decrease in the temperature of the gas.
- Bursting of a cycle tube.
- Propagation of sound waves in a gaseous medium.
- In diesel engines burning of diesel without spark plug is done due to adiabatic compression of diesel vapour and air mixture

Free Expansion

Take a thermally insulated bottle with ideal gas at some temperature T_1 and, by means of a pipe with a stopcock, connect this to another insulated bottle which is evacuated. If we suddenly open the stopcock, the gas will rush from the first bottle into the second until the pressures are equalized.



free expansion of a gas

Experimentally, we find that this process of free expansion does not change the temperature of the gas – when the gas attains equilibrium and stops flowing, the final temperature of both bottles are equal to the initial temperature T_1 . This process is called a free expansion.

The change in the internal energy of the gas can be calculated by applying the first law of thermodynamics to the free-expansion process.

The process is adiabatic because of the insulation, So $Q = 0$.

No part of the surroundings moves so the system does no work on its surroundings.

- **For ideal gas** $(\delta W)_{\text{ext.}} = \text{Work done against external atmosphere} = P dV = 0$
 $(\because P = 0) \Rightarrow \text{because vacuum}$
 $(\delta W)_{\text{int.}} = \text{Work done against internal molecular forces} = 0$
 $\delta Q = dU + \delta W \Rightarrow 0 = dU + 0$

The internal energy does not change $dU = 0$ So U and T are constant.

The initial and final states of this gas have the same internal energy.

Which implies that the internal energy of an ideal gas does not depend on the volume at all.

The free-expansion process has led us to the following conclusion :

The internal energy $U(T)$ of an ideal gas depends only on the temperature.

Relation between degrees of freedom and specific heat of gas

Energy related with each degree of freedom for a gas molecule = $\frac{1}{2} kT$, Energy related with all degree of freedom for a gas molecule = $\frac{f}{2} kT$

Internal energy of one mole of ideal gas (total K.E.) $U = \frac{f}{2} kT$ for Isometric process (volume constant) $\delta W = 0$

By first law of thermodynamics $\delta Q = \delta W + dU \Rightarrow C_V dT = dU \Rightarrow C_V = \frac{dU}{dT}$

$$C_V = \frac{dU}{dT} = \frac{f}{2} R = \frac{R}{\gamma - 1}, C_P = C_V + R = \left[\frac{f}{2} + 1 \right] R = \frac{\gamma R}{\gamma - 1} \text{ and } \gamma = \frac{C_P}{C_V} = 1 + \frac{2}{f}$$

$$C_V = \frac{R}{\gamma - 1}, C_P = \frac{\gamma R}{\gamma - 1} \text{ and } \gamma = 1 + \frac{2}{f}$$

General expression for C (C_P or C_V) in the process $PV^x = \text{constant}$ $C = \frac{R}{\gamma - 1} + \frac{R}{1 - x}$

For isobaric process $P = \text{constant}$ so $x = 0 \therefore C = C_P = \frac{R}{\gamma - 1} + R = C_V + R$

For isothermal process, $PV = \text{constant}$ so $x = 1 \therefore C = \infty$

For adiabatic process $PV^\gamma = \text{constant}$ so $x = \gamma \therefore C = 0$

Values of f, U, C_V , C_P and γ for different gases are shown in table below.

Atomicity of gas	f	C_V	C_P	γ
Monoatomic	3	$\frac{3}{2}R$	$\frac{5}{2}R$	$\frac{5}{3} = 1.67$
Diatomic, Triatomic and Triatomic linear (at normal temperature)	5	$\frac{5}{2}R$	$\frac{7}{2}R$	$\frac{7}{5} = 1.4$
Poly atomic Triangular Non-linear	6	$\frac{6}{2}R = 3R$	$\frac{8}{2}R = 4R$	$\frac{4}{3} = 1.33$

- $1 < \gamma < 2$
- If atomicity of gases is same U, C_P , C_V and γ is same for gas mixture.
- If in a gas mixture gases are of different atomicity, then for gas mixture γ changes according to following condition. Diatomic $\gamma_1 \leq \text{mixture} \leq \gamma_2$ mono atomic where $\gamma_1 < \gamma_2$
- If 'f' is the degree of freedom per molecule for a gas, then

* Total energy of each molecule = $\frac{fkT}{2}$

* Total energy per mole of gas = $N_0 \frac{f}{2} kT = \frac{f}{2} RT$

- According to kinetic theory of gases, the molecule are not interacting with each other. So potential energy is zero and internal energy of gas molecules is only their kinetic energy.
- For ' μ ' mole of a gas : Internal energy at temperature T is $U = \frac{\mu f R T}{2} = \mu C_V T$
- Change in internal energy is given by $dU = \frac{\mu f R}{2} (dT) = \mu C_V dT$. This change is process independent.

C_p is greater than C_v

If a gas is heated at constant volume, the gas does no work against external pressure. In this case, the whole of the heat energy supplied to the gas is spent in raising the temperature of the gas.

If a gas is heated at constant pressure, its volume increases. In this case, heat energy is required for the following two purposes :

- to increase the volume of the gas against external pressure.
- To increase the temperature of 1 mole of gas through 1 K.

Thus, more heat energy is required to raise the temperature of 1 mole of gas through 1 K when it is heated at constant pressure than when it is heated at constant volume. $\therefore C_p > C_v$

The difference between C_p and C_v is equal to thermal equivalent of the work done by the gas in expanding against external pressure.

Mayer's formula : $C_p - C_v = R$

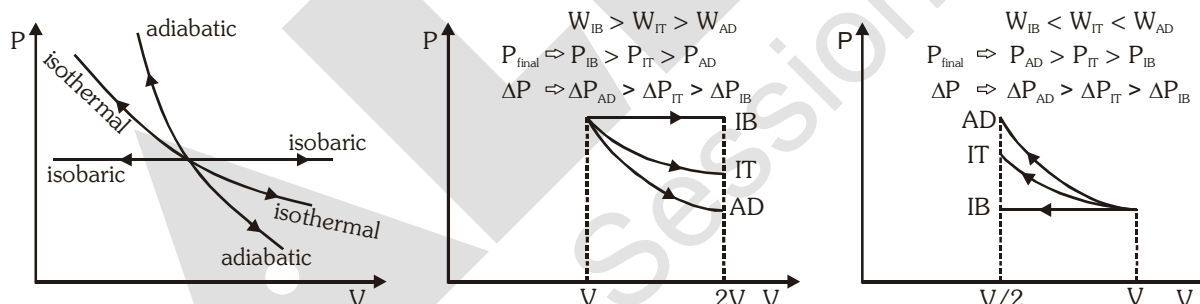
\therefore At constant pressure $dQ = \mu C_p dT$, $dU = \mu C_v dT$ & $dW = PdV = \mu R dT$

Now from first law of thermodynamics $dQ = dW + dU$

$$\Rightarrow \mu C_p dT = \mu R dT + \mu C_v dT \Rightarrow C_p = R + C_v \Rightarrow C_p - C_v = R$$

GOLDEN KEY POINTS

- When a gas expands its volume increases, then final pressure is less for adiabatic expansion. But, when a gas compresses its volume decreases, then the final pressure is more in case of adiabatic compression.

**First Law of Thermodynamics Applied to Different Processes**

Process	Q	ΔU	W
Cyclic	W	0	Area of the closed curve
Isochoric	ΔU	$\mu C_v \Delta T$ (μ mole of gas)	0
Isothermal	W	0	$\mu RT \log_e \left[\frac{V_f}{V_i} \right] = \mu RT \log_e \left[\frac{P_i}{P_f} \right]$
Adiabatic	0	-W	$\frac{\mu R(T_f - T_i)}{1 - \gamma}$
Isobaric	$\mu C_p \Delta T$	$\mu C_v \Delta T$	$P(V_f - V_i) = \mu R(T_f - T_i)$

Illustrations

Illustration 56.

An ideal gas has a specific heat or molar heat capacity at constant pressure $C_p = 5R/2$. The gas is kept in a closed vessel of volume 0.0083 m^3 at a temperature of 300K and a pressure of $1.6 \times 10^6 \text{ Nm}^{-2}$. An amount of $2.49 \times 10^4 \text{ J}$ of heat energy is supplied to the gas. Calculate the final temperature and pressure of the gas.

Solution

$$C_v = C_p - R = \frac{5R}{2} - R = \frac{3R}{2}, \Delta V = 0, T_1 = 300 \text{ K}, V = 0.0083 \text{ m}^3, P_1 = 1.6 \times 10^6 \text{ Nm}^{-2}$$

$$\text{From first law of thermodynamics } Q = \Delta U + P\Delta V \Rightarrow \Delta U = Q = 2.49 \times 10^4 \text{ J}$$

$$\text{From gas equation } n = \frac{PV}{RT} = \frac{1.6 \times 10^6 \times 0.0083}{8.3 \times 300} = \frac{16}{3} \therefore \Delta U = nC_v\Delta T \Rightarrow \Delta T = \frac{\Delta U}{nC_v} = \frac{2.49 \times 10^4 \times 6}{3 \times 8.3 \times 16} = 375 \text{ K}$$

$$\text{Final temperature} = 300 + 375 = 675\text{K}$$

$$\text{According to pressure law } P \propto T \Rightarrow \frac{P_2}{P_1} = \frac{T_2}{T_1} \Rightarrow P_2 = \frac{T_2}{T_1} \times P_1 = \frac{1.6 \times 10^6 \times 675}{300} = 3.6 \times 10^6 \text{ Nm}^{-2}$$

Illustration 57.

5 moles of oxygen is heated at constant volume from 10°C to 20°C . What will be change in the internal energy of the gas? The gram molecular specific heat of oxygen at constant pressure.

$$C_p = 8 \text{ cal/mole } \text{ and } R = 8.36 \text{ joule/mole } ^\circ\text{C}.$$

Solution

$$\therefore C_v = C_p - R = 8 - 2 = 6 \text{ cal/mole } ^\circ\text{C}$$

$$\therefore \text{Heat absorbed by 5 moles of oxygen at constant volume}$$

$$Q = \mu C_v \Delta T = 5 \times 6 (20 - 10) = 30 \times 10 = 300 \text{ cal}$$

$$\text{At constant volume } \Delta V = 0. \therefore \Delta W = 0$$

$$\therefore \text{From first law of thermodynamics } Q = \Delta U + W \Rightarrow 300 = \Delta U + 0 \Rightarrow \Delta U = 300 \text{ cal}$$

Illustration 58.

What amount of heat must be supplied to $2.0 \times 10^{-2} \text{ kg}$ of nitrogen (at room temperature) to raise the temperature by 45°C at constant pressure. Molecular mass of $\text{N}_2 = 28$, $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$.

Solution

$$\text{Here } m = 2 \times 10^{-2} \text{ kg}, \Rightarrow \mu = \frac{M}{M_w} = \frac{2 \times 10^{-2}}{28 \times 10^{-3}} = \frac{5}{7} \text{ \& } C_p = \frac{7}{2} R \therefore Q = \mu C_p \Delta T = \frac{5}{7} \times \frac{7}{2} \times 8.3 \times 45 = 933.75 \text{ J}$$

Illustration 59.

Two moles of a gas at 127°C expand isothermally until its volume is doubled. Calculate the amount of work done.

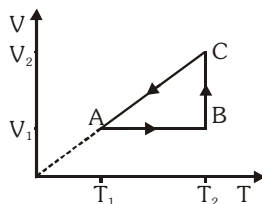
Solution

$$\mu = 2, \quad T = 127 + 273 = 400\text{K}, \quad \frac{V_2}{V_1} = 2$$

$$\begin{aligned} \text{From formula } W &= 2.3026 \mu RT \log_{10} \frac{V_2}{V_1} = 2.3026 \times 2 \times 8.3 \times 400 \times \log_{10} 2 \\ &= 2.3026 \times 2 \times 8.3 \times 400 \times 0.3010 \approx 4.6 \times 10^3 \text{ J} \end{aligned}$$

Illustration 60.

Figure shows a process ABCA performed on an ideal gas. Find the net heat given to the system during the process.

**Solution**

Since the process is cyclic, hence the change in internal energy is zero.

The heat given to the system is then equal to the work done by it.

The work done in part AB is $W_1 = 0$ (the volume remains constant). The part BC represents an isothermal process so that the work done by the gas during this part is $W_2 = \mu RT_2 \ln \frac{V_2}{V_1}$

During the part CA : $V \propto T$ So, V/T is constant and hence, $P = \frac{\mu RT}{V}$ is constant

The work done by the gas during the part CA is $W_3 = P(V_1 - V_2) = \mu RT_1 - \mu RT_2 = -\mu R (T_2 - T_1)$.

The net work done by the gas in the process ABCA is $W = W_1 + W_2 + W_3 = \mu R \left[T_2 \ln \frac{V_2}{V_1} - (T_2 - T_1) \right]$

The same amount of heat is given to the gas.

Illustration 61.

During an experiment an ideal gas is found to obey an additional law $VP^2 = \text{constant}$. The gas is initially at temperature T and volume V . What will be the temperature of the gas when it expands to a volume $2V$.

Solution

By gas equation $PV = \mu RT$ and $VP^2 = \text{constant}$ on eliminating P

$$\left[\frac{A}{\sqrt{V}} \right] V = \mu RT \Rightarrow \sqrt{V} = \frac{\mu R}{A} T \quad \therefore \frac{\sqrt{V_1}}{\sqrt{V_2}} = \left[\frac{T_1}{T_2} \right] \Rightarrow \frac{\sqrt{V}}{\sqrt{2V}} = \frac{T}{T'} \Rightarrow T' = (\sqrt{2}) T$$

Illustration 62.

2m^3 volume of a gas at a pressure of $4 \times 10^5 \text{ Nm}^{-2}$ is compressed adiabatically so that its volume becomes 0.5m^3 . Find the new pressure. Compare this with the pressure that would result if the compression was isothermal.

Calculate work done in each process. $\gamma = 1.4$

Solution

$$V_1 = 2\text{m}^3, P_1 = 4 \times 10^5 \text{ Nm}^{-2}, \quad V_2 = 0.5\text{m}^3$$

$$\text{In adiabatic process} \quad P_1 V_1^\gamma = P_2 V_2^\gamma \Rightarrow P_2 = 4 \times 10^5 \left[\frac{2}{0.5} \right]^{1.4} = 4 \times 10^5 (4)^{1.4} = 2.8 \times 10^6 \text{ Nm}^{-2}$$

$$\text{In isothermal process} \quad P_1 V_1 = P_2 V_2 \Rightarrow P_2 = \frac{P_1 V_1}{V_2} = \frac{4 \times 10^5 \times 2}{0.5} = 1.6 \times 10^6 \text{ Nm}^{-2}.$$

$$\text{Now work done in adiabatic process} \quad W = \frac{P_1 V_1 - P_2 V_2}{\gamma - 1} = \frac{(4 \times 10^5 \times 2 - 2.8 \times 10^6 \times 0.5)}{1.4 - 1} = -1.48 \times 10^6 \text{ J}$$

$$\begin{aligned} \text{Work done in isothermal process} \quad W &= 2.303 \mu RT \log \frac{V_2}{V_1} = 2.3026 P_1 V_1 \log \frac{V_2}{V_1} \\ &= 2.3026 \times 4 \times 10^5 \times 2 \times \log \left[\frac{0.5}{2.0} \right] = 2.3026 \times 4 \times 10^5 \times 2 \log \left(\frac{1}{4} \right) = -1.1 \times 10^6 \text{ J} \end{aligned}$$

Illustration 63.

Two samples of a gas initially at same temperature and pressure are compressed from a volume V to $\frac{V}{2}$. One sample is compressed isothermally and the other adiabatically. In which sample is the pressure greater?

Solution

Let initial volume, $V_1 = V$ and pressure, $P_1 = P$, final volume, $V_2 = \frac{V}{2}$ and final pressure, $P_2 = ?$

For isothermal compression $P_2 V_2 = P_1 V_1$ or $P_2 = \frac{P_1 V_1}{V_2} = \frac{PV}{\frac{V}{2}} = 2P$

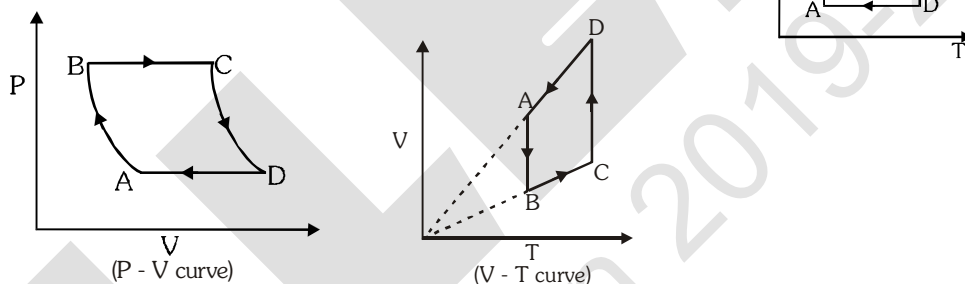
For adiabatic compression $P_2' = P_1 \left[\frac{V_1}{V_2} \right]^\gamma \Rightarrow P_2' = P \left[\frac{V}{V/2} \right]^\gamma = 2^\gamma P \Rightarrow P_2' = 2^\gamma P$ $\gamma > 1 \therefore 2^\gamma > 2$ and $P_2' > P_2$

Pressure during adiabatic compression is greater than the pressure during isothermal compression.

Illustration 64.

Plot $P - V$, $V - T$ graph corresponding to the $P - T$ graph for an ideal gas shown in figure. Explain your answers.

Solution



For process AB ; $T = \text{constant}$ so $P \propto \frac{1}{V}$

For process BC $P = \text{constant}$ so $V \propto T$

For process CD ; $T = \text{constant}$ so $V \propto \frac{1}{P}$

For process DA $P = \text{constant}$ so $V \propto T$

Illustration 65.

The specific heat of argon at constant volume is 0.075 kcal/kg K . Calculate its atomic weight, $[R = 2 \text{ cal/mol K}]$

Solution

As argon is monoatomic, its molar specific heat at constant volume will be

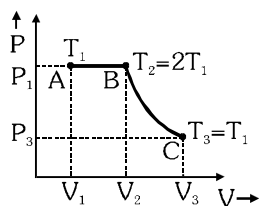
$$C_v = \frac{3}{2}R = \frac{3}{2} \times 2 = 3 \text{ cal/mol K}, C_v = M_w c_v \text{ and } c_v = 0.075 \text{ cal/g K}$$

$$\text{So } 3 = M_w \times 0.075 \Rightarrow M_w = \frac{3}{0.075} = 40 \text{ gram/mole}$$

Illustration 66.

Two moles of helium gas ($\gamma = 5/3$), assumed ideal, are initially at 27°C and occupy a volume of 20 litres. The gas is first expanded at constant pressure till its volume is doubled. It then undergoes an adiabatic change until the temperature returns to its initial value. Determine the final pressure and volume of the gas. Also, calculate the work done under isobaric and adiabatic processes. $[R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}]$ [AIPMT (Mains) 2005]

Solution



$$V_1 = 20 \times 10^{-3} \text{ m}^3,$$

$$T_3 = T_1 = 300\text{K}$$

For process AB ($P = \text{constant}$) $T_2 = 2T_1 = 600\text{K} (\because V_2 = 2V_1)$

For process BC $TV^{\gamma-1} = \text{constant}$ Therefore $(2T_1)(V_2)^{5/3-1} = (T_1)(V_3)^{5/3-1}$

$$\Rightarrow V_3 = (V_2)(2)^{3/2} = 2\sqrt{2}V_2 = 80\sqrt{2} \times 10^{-3} \text{ m}^3 = 80\sqrt{2} \text{ litres}$$

$$P_3 = \frac{\mu RT_3}{V_3} = \frac{(2)(8.3)(300)}{80\sqrt{2} \times 10^{-3}} = 0.44 \times 10^5 \text{ N/m}^2$$

Work done under isobaric process = $P\Delta V = P_1(V_2 - V_1)$

$$= \frac{\mu RT_1}{V_1}(V_2 - V_1) = \frac{2 \times 8.3 \times 300}{20 \times 10^{-3}}(40 \times 10^{-3} - 20 \times 10^{-3}) = 4980\text{J}$$

$$\text{Work done under adiabatic process} = \frac{\mu R(T_2 - T_3)}{(\gamma - 1)} = \frac{2 \times 8.3 \times (300)}{(5/3 - 1)} = 7470\text{J}$$

BEGINNER'S BOX-8

- 32 gram of oxygen gas at temperature 27°C is compressed adiabatically to $1/3$ of its initial volume. Calculate the change in internal energy? ($\gamma = 1.5$ for oxygen)
- How much heat energy should be added to the gaseous mixture consisting of one gram of hydrogen and one gram helium to raise its temperature from 0°C to 50°C .
 - at constant volume and
 - at constant pressure, assuming both the gases to be ideal.

$$R = 2 \text{ cal/mol-K}, \gamma_{\text{H}_2} = 1.41, \gamma_{\text{He}} = 1.67$$

3. Match the column-I with column-II

Column - I	Column - II
(A)	(p) Isometric process
(B)	(q) Adiabatic process
(C)	(r) Isothermal process

4. Match the column-I with column-II

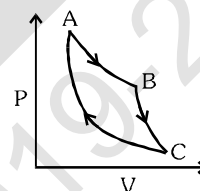
Column-I	Column-II
(A) Isobaric process	(p) No heat exchange
(B) Isothermal process	(q) Constant pressure
(C) Isoentropic process	(r) Constant internal energy
(D) Isochoric process	(s) Work done is zero

5. Match the column-I with column-II

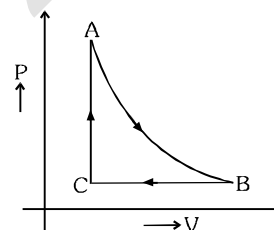
Column-I	Column-II
(A) Adiabatic expansion	(p) $Q = 0$
(B) Adiabatic free expansion	(q) $\Delta U = 0$
(C) Isobaric expansion	(r) $W = 0$
(D) Isothermal expansion	(s) $Q = W + \Delta U$
(E) Isochoric cooling	

6. An ideal gas expands isothermally along AB and does 700 J of work.

- How much heat does the gas exchange along AB.
- The gas expands adiabatically along BC and does 400 J of work. When the gas returns to A along CA, it exhusts 100 J of heat to its surroundings. How much work is done on the gas along this path.



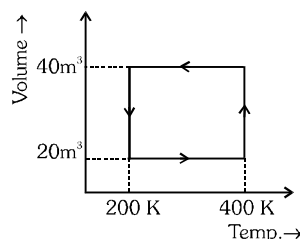
7. For one mole of an ideal gas at 300 K occupies a volume of 0.36 m^3 at a pressure of 2 atm. The gas expands adiabatically until its volume becomes 1.44 m^3 . Next gas is compressed isobarically to its original volume. Finally, the pressure is increased isochorically until the gas returns to its initial state. (see figure).



- Determine the temp. at the point B and C
- The work done during the process A to B

(Assume $\gamma = 1.5$, $R = 8 \frac{\text{J}}{\text{mol} \cdot \text{K}}$)

8. Following figure is given for 1 mole gas.



Then find out net work done in the cyclic process.

EFFICIENCY OF A CYCLE (η) :

$$\eta = \frac{\text{total Mechanical work done by the gas in the whole process}}{\text{Heat absorbed by the gas (only +ve)}}$$

$$= \frac{\text{area of the P-V cycle}}{\text{Heat injected into the system}}$$

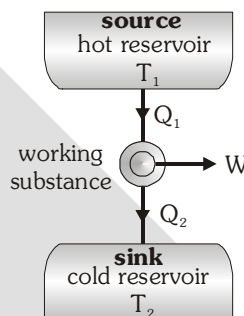
HEAT ENGINE

Heat engine is a device which converts heat into work.

Three parts of a heat engine:

- (i) Source of high temperature reservoir at temperature T_1
- (ii) Sink of low temperature reservoir at temperature T_2
- (iii) Working substance.

In a cycle of heat engine the working substance extracts heat Q_1 from source, does some work W and rejects remaining heat Q_2 to the sink.



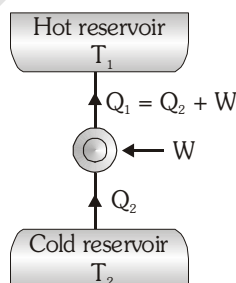
$$\text{Efficiency of heat engine } \eta = \frac{\text{work done (W)}}{\text{heat taken from source (Q}_1\text{)}} \Rightarrow \frac{T_1 - T_2}{T_1} = \frac{Q_1 - Q_2}{Q_1}$$

REFRIGERATOR

It is inverse of heat engine. It extracts heat (Q_2) from a cold reservoir, some external work W is done on it and rejects heat (Q_1) to hot reservoir.

The coefficient of performance of a refrigerator (C.O.P. or β).

$$\beta = \frac{\text{Heat extracted from cold reservoir}}{\text{Work done on refrigerator}} = \frac{Q_2}{W} = \frac{Q_2}{Q_1 - Q_2} = \frac{1}{\frac{Q_1}{Q_2} - 1}$$



$$\text{For Carnot reversible refrigerator } \frac{Q_1}{Q_2} = \frac{T_1}{T_2} \therefore \beta = \frac{Q_2}{W} = \frac{1}{\left[\frac{Q_1}{Q_2} - 1\right]} = \frac{1}{\left[\frac{T_1}{T_2} - 1\right]}$$

$$\Rightarrow \beta = \frac{T_2}{T_1 - T_2}$$

Relation between β and η :

$$\therefore \eta = \frac{W}{Q_1} \quad [\text{Efficiency of heat engine}]$$

(Refrigerator and heat engine are maintained at same thermal reservoirs)

$$\beta = \frac{Q_2}{W} \quad [\text{C.O.P. of refrigerator}]$$

$$\text{and } Q_1 = Q_2 + W$$

$$\text{then } \boxed{\beta = \frac{1}{\eta} - 1}$$

SECOND LAW OF THERMODYNAMICS

The first law of thermodynamics is the principle of conservation of energy. Common experience shows that there are many conceivable processes that are perfectly allowed by the First law and yet are never observed. For example, nobody has ever seen a book lying on a table jumping to a height by itself. Would be possible if the principle of conservation of energy were the only restriction. This principle, which disallows many phenomena consistent with the first law of thermodynamics is known as the second law of thermodynamics. The Second Law of Thermodynamics gives a fundamental limitation to the efficiency of a heat engine and the co-efficient of performance of refrigerator.

Second Law of Thermodynamics

Kelvin Planck Statement

No process is possible whose sole result is the absorption of heat from a reservoir and the complete conversion of the heat into work.

Claussius Statement

No process is possible whose sole result is the transfer of heat from a colder object to a hotter object.

It can be proved that the two statements above are completely equivalent.

The second law of thermodynamics is a general principle which places constraints upon the direction of heat transfer and the attainable efficiencies of heat engines.

Entropy Statement of the second Law of Thermodynamics

The entropy of the Universe never decrease. It either increases (for irreversible processes) or remains the same (for reversible processes)

REVERSIBLE AND IRREVERSIBLE PROCESS

(i) Reversible Process

A reversible process is defined as a process that can be reversed without leaving any trace on the surroundings. Both the system and the surroundings are returned to their initial states at the end of the reverse process.

Conditions for a Reversible process

(a) The process should take place very slowly, i.e. quasi-statically, i.e. seemingly static (from the Latin word qasi meaning 'as if') so that it satisfies the following requirements at each stage of the process.

- The system should be in mechanical equilibrium.
- The system should be in thermal equilibrium.
- The system should be in chemical equilibrium, i.e., no new products should be formed.

(b) There should be no friction losses etc. It should be remembered that a complete reversible process or cycle of operations is only an ideal case. In actual practice, there is always a loss of heat due to friction, conduction and radiation. However, it is possible to approximate reversible processes through carefully controlled procedures, but they can never be achieved.

Examples

- Frictionless pendulum.
- Quasi-equilibrium expansion and compression of a gas.

(b) Irreversible Process

A process that is not reversible is called an irreversible process.

The spontaneous processes occurring in nature are irreversible. In fact irreversibility is a rule rather than an exception in nature.

Examples

- (i) Sudden unbalanced expansion of a gas.
- (ii) Heat produced by friction.
- (iii) Heat generated when a current flows through an electric resistance in any direction.
- (iv) Cooking gas leaking from a cylinder.
- (v) Diffusion of liquids or gases.
- (vi) Breaking of an egg.
- (vii) The growth of a plant.

A system can be restored to its initial state following a process, regardless of whether the process is reversible or irreversible. But for reversible process, this restoration is made without leaving any net change on the surroundings whereas for irreversible processes, the surroundings usually do some work on the system and therefore will not return to their original state.

ENTROPY AND THE SECOND LAW OF THERMODYNAMICS

To have a better understanding of thermodynamic process and the second and third laws of thermodynamics, let us introduce the concept of entropy. It was introduced by Clausius in 1865.

A quantity that denotes the amount of disorder is called entropy and is denoted by S .

The total energy always conserved in any process but the total entropy always increases or remains the same in any process (ie., the disorder increases or remain constant)

$$dS = \frac{dQ}{T} \Rightarrow \Delta S = \int \frac{dQ}{T} \quad \dots(1)$$

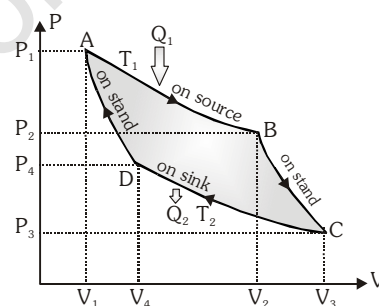
For reversible processes, $\Delta S = 0$. For irreversible process $\Delta S > 0$

A process where $\Delta S < 0$ is not possible.

3. Carnot cycle

Carnot devised an ideal engine which is based on a reversible cycle of four operations in succession :

- (i) Isothermal expansion, $A \rightarrow B$
- (ii) Adiabatic expansion, $B \rightarrow C$
- (iii) Isothermal compression $C \rightarrow D$
- (iv) Adiabatic compression. $D \rightarrow A$

**Main parts of Carnot's engine are**

- **Source of heat :**

It is a hot body of very large heat capacity kept at a constant high temperature T_1 . Its upper surface is perfectly conducting so that working substance can take heat from it.

- **Mechanical arrangements and working substance :**

It is a hollow cylinder whose walls are perfectly non-conducting and its base is perfectly conducting fitted with non-conducting piston. This piston move without any friction. Ideal gas enclosed in cylinder as a working substance.

- **Heat sink :**

It is a cold body at low temperature T_2 . It is a body of large heat capacity and its upper surface is highly conducting so that working substance can reject heat to it.

● **Stand :**

It is made by perfectly insulating material. When cylinder placed on it working substance can expanded or compressed adiabatic.

● **Working :**

A set of reversible processes through them working substance is taken back to initial condition to get maximum work from this type of ideal engine.

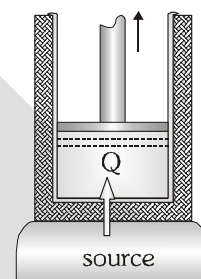
Processes of Carnot's cycle can be denoted by an indicator diagram.

● **Isothermal expansion A → B**

Initially the cylinder is taken to be in thermal equilibrium with the high temperature T_1 , this is initial state of working substance denoted by point A (P_1, V_1, T_1). After that the piston is allowed to move outward slowly. With the movement of the piston. The process is very slow so that it is isothermal. Heat from reservoir flows through the base of cylinder into the gas so temperature of the gas remains T_1 . Gas expand and receive heat Q_1 from source and gets state B (P_2, V_2, T_1). This heat input Q_1 to the gas from path A to B is utilized for doing work W_1 .

By path A to B the heat input to the gas = the work done against the external pressure.

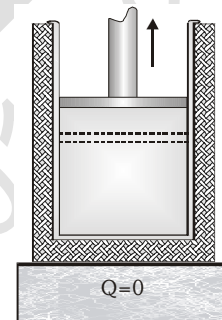
$$W_1 = Q_1 = \int_{V_1}^{V_2} P dV = \int_{V_1}^{V_2} \frac{\mu R T_1}{V} dV = \mu R T_1 \ln \frac{V_2}{V_1}$$



● **Adiabatic expansion B → C**

Now cylinder is put in contact with a non-conducting stand and piston is allowed to move outward, because no heat can enter in or leave out so the expansion of gas is adiabatic. The temperature falls to T_2 K and gas describes the adiabatic from B to point C (P_3, V_3, T_2). During this expansion more work is done (W_2) at the expense of the internal energy.

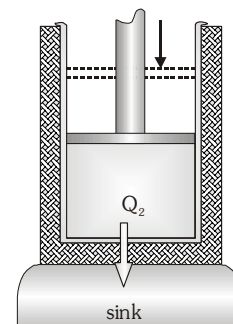
$$\text{Work done in adiabatic path BC is } W_2 = \frac{\mu R}{\gamma - 1} (T_1 - T_2)$$



● **Isothermal compression C → D**

Now the gas cylinder is placed in contact with sink at temperature T_2 . The piston is moved slowly inward so that heat produced during compression passes to the sink. The gas is isothermally compressed from C to point D (P_4, V_4, T_2). The heat rejected Q_2 to the cold reservoir (sink) at T_2 occurs over this path. Amount of work done on gas W_3 = amount of heat rejected to the sink

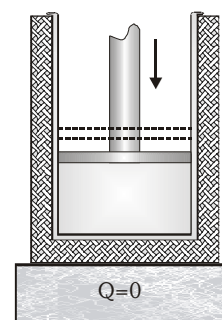
$$Q_2 = W_3 = \mu R T_2 \ln \left(\frac{V_4}{V_3} \right) \Rightarrow Q_2 = \mu R T_2 \ln \left(\frac{V_4}{V_3} \right)$$



● **Adiabatic compression D → A**

The cylinder is removed from the sink and is put in contact with insulating stand now piston moves inward. Heat is not allowed to go out and it increases the internal energy of the system. Now work is done on the gas during adiabatic compression from state D to initial point A (P_1, V_1, T_1). No heat exchanges occur over the adiabatic path. Work done on the system $W_4 = \frac{\mu R}{\gamma - 1} (T_2 - T_1)$

$$\text{This cycle of operations is called a Carnot cycle.}$$



In first two steps work is done by engine W_1 and W_2 are positive

In last two steps work is done on gas W_3 and W_4 are negative

The work done in complete cycle W = the area of the closed part of the P-V cycle.

$$W = W_1 + W_2 + W_3 + W_4$$

$$\therefore W = \mu RT_1 \ln \frac{V_2}{V_1} + \frac{\mu R}{\gamma - 1} (T_1 - T_2) + \mu RT_2 \ln \frac{V_4}{V_3} - \frac{\mu R}{\gamma - 1} (T_2 - T_1) = \mu RT_1 \ln \frac{V_2}{V_1} + \mu RT_2 \ln \frac{V_4}{V_3}$$

$$\text{Efficiency of Carnot Engine, } \eta = \frac{W}{Q_1} \Rightarrow \eta = \frac{\mu RT_1 \ln \frac{V_2}{V_1} + \mu RT_2 \ln \frac{V_4}{V_3}}{\mu RT_1 \ln \frac{V_2}{V_1}}$$

B to C and D to A are adiabatic paths

$$\text{so } T_1 V_2^{(\gamma-1)} = T_2 V_3^{(\gamma-1)} \text{ and } T_1 V_1^{(\gamma-1)} = T_2 V_4^{(\gamma-1)} \Rightarrow \frac{V_2}{V_1} = \frac{V_3}{V_4}$$

$$\Rightarrow \eta = \frac{T_1 - T_2}{T_1} = \frac{Q_1 - Q_2}{Q_1} \Rightarrow \eta = 1 - \frac{Q_2}{Q_1} = 1 - \frac{T_2}{T_1} \frac{Q_1}{T_1} = \frac{Q_2}{T_2}$$

$$\eta = \frac{T_1 - T_2}{T_1} \times 100\% \Rightarrow \eta = \frac{Q_1 - Q_2}{Q_1} \times 100\%$$

The efficiency for the Carnot engine is the best that can be obtained for any heat engine. The efficiency of a Carnot engine is never 100% because it is 100% only if temperature of sink $T_2 = 0$ which is impossible.

Carnot Theorem

No irreversible engine (I) can have efficiency greater than Carnot reversible engine (R) working between same

hot and cold reservoirs. $\eta_R > \eta_I \Rightarrow 1 - \frac{T_2}{T_1} > 1 - \frac{Q_2}{Q_1}$

Illustrations

Illustration 67.

A carnot engine working between 400 K and 800 K has a work output of 1200 J per cycle. What is the amount of heat energy supplied to the engine from source per cycle?

Solution

$$W = 1200 \text{ J, } T_1 = 800 \text{ K, } T_2 = 400 \text{ K} \therefore \eta = 1 - \frac{T_2}{T_1} = \frac{W}{Q_1} \Rightarrow 1 - \frac{400}{800} = \frac{1200}{Q_1}$$

$$\Rightarrow 0.5 = \frac{1200}{Q_1}$$

$$\text{Heat energy supplied by source } Q_1 = \frac{1200}{0.5} = 2400 \text{ joule per cycle}$$

Illustration 68.

The temperatures T_1 and T_2 of the two heat reservoirs in an ideal carnot engine are 1500°C and 500°C respectively. Which of the following : increasing T_1 by 100°C or decreasing T_2 by 100°C would result in a greater improvement in the efficiency of the engine?

Solution

$$T_1 = 1500^\circ\text{C} = 1500 + 273 = 1773 \text{ K and } T_2 = 500^\circ\text{C} = 500 + 273 = 773 \text{ K.}$$

$$\text{The efficiency of a carnot's engine } \eta = 1 - \frac{T_2}{T_1}$$

When the temperature of the source is increased by 100°C , keeping T_2 unchanged, the new temperature of the source is $T'_1 = 1500 + 100 = 1600^\circ\text{C} = 1873 \text{ K}$.

The efficiency becomes $\eta' = 1 - \frac{T_2}{T'_1} = 1 - \frac{773}{1873} = 0.59$

On the other hand, if the temperature of the sink is decreased by 100°C , keeping T_1 unchanged, the new temperature of the sink is $T'_2 = 500 - 100 = 400^\circ\text{C} = 673 \text{ K}$. The efficiency now becomes

$$\eta'' = 1 - \frac{T'_2}{T_1} = 1 - \frac{673}{1773} = 0.62$$

Since η'' is greater than η' , decreasing the temperature of the sink by 100°C results in a greater efficiency than increasing the temperature of the source by 100°C .

Illustration 69.

A heat engine operates between a cold reservoir at temperature $T_2 = 300 \text{ K}$ and a hot reservoir at temperature T_1 . It takes 200 J of heat from the hot reservoir and delivers 120 J of heat to the cold reservoir in a cycle. What could be the minimum temperature of hot reservoir?

Solution

Work done by the engine in a cycle is $W = 200 - 120 = 80 \text{ J}$. $\eta = \frac{W}{Q} = \frac{80}{200} = 0.4$

From Carnot's Theorem $0.4 \leq 1 - \frac{T_2}{T_1} = 1 - \frac{300}{T_1} \Rightarrow \frac{300}{T_1} \leq 0.6 \Rightarrow T_1 \geq \frac{300}{0.6} \Rightarrow T_1 \geq 500$

Illustration 70.

A Carnot engine works as a refrigerator between 250 K and 300 K . If it receives 750 cal of heat from the reservoir at the lower temperature. Calculate the amount of heat rejected at the higher temperature.

Solution

$T_1 = 300 \text{ K}$ $T_2 = 250 \text{ K}$ $Q_2 = 750$ $Q_1 = ?$

$$\frac{Q_1}{Q_2} = \frac{T_1}{T_2} \quad Q_1 = \frac{300}{250} \times 750 = 900 \text{ cal}$$

Illustration 71.

The temperature inside and outside of refrigerator are 260 K and 315 K respectively. Assuming that the refrigerator cycle is reversible, calculate the heat delivered to surroundings for every joule of work done.

Solution

$T_2 = 260 \text{ K}$, $T_1 = 315 \text{ K}$; $W = 1 \text{ joule}$

Coefficient of performance of Carnot refrigerator $\beta = \frac{Q_2}{W} = \frac{T_2}{T_1 - T_2}$

$$\therefore \frac{Q_2}{1} = \frac{260}{315 - 260} = \frac{260}{55} \Rightarrow Q_2 = \frac{260}{55} = 4.73 \text{ J} \quad \text{So} \quad Q_1 = Q_2 + W = 5.73 \text{ J}$$

Illustration 72.

A refrigerator takes heat from water at 0°C and transfer it to room at 27°C . If 100 kg of water is converted in ice at 0°C then calculate the work done. (Latent heat of ice is $3.4 \times 10^5 \text{ J/kg}$)

Solution

$$\text{Coefficient of performance (COP)} = \frac{T_2}{T_1 - T_2} = \frac{273}{300 - 273} = \frac{273}{27}$$

$$W = \frac{Q_2}{\text{COP}} = \frac{mL}{\text{COP}} = \frac{100 \times 3.4 \times 10^5}{273/27} = \frac{100 \times 3.4 \times 10^5 \times 27}{273} = 3.36 \times 10^6 \text{ J}$$

BEGINNER'S BOX-9

1. An ideal engine which works between 27°C & 227°C takes 100 calorie of heat in one cycle. Calculate the work done in one cycle ?
2. The coefficient of performance of a carnot refrigerator is $5/3$. If the refrigerator operating at a intake temperature -23°C then what will be the exhaust temperature ?
3. A carnot engine absorbs 1000 J of heat from a reservoir at 127°C and reject 600 J of heat during each cycle.

Calculate

- (a) The efficiency of engine
- (b) The temperature of the sink
- (c) The amount of the useful work done during each cycle.

4. Two carnot engine A and B use one after another. Engine A absorb heat from source at $T_1 = 800 \text{ K}$ and release heat to sink of temp $T_2 \text{ K}$. Second engine B absorb heat which release by first (at $T_2 \text{ K}$) and release heat at another sink at temperature $T_3 = 300 \text{ K}$. If work done by two engine is same than calculate value of T_2 ?
5. A carnot engine, efficiency 40% and temperature of sink 300 K to increase efficiency up to 60% calculate change in temperature of source ?
6. An ideal refrigerator is working between temperature 27°C and 127°C . If it expells 120 calorie of heat in one second then calculate its wattage ?

ANSWERS

BEGINNER'S BOX-1

1. $50\text{ K} < 50^\circ\text{F} < 50^\circ\text{C}$
2. (a) All tie (b) 50°X , 50°Y , 50°W .
3. -40°C or -40°F .
4. (B) 5. $1.44 \times 10^{-2}\text{ cm}$
6. In bimetallic strips the two metals have different thermal expansion coefficient. Hence on heating it bends towards the metal with lower thermal expansion coefficient.

BEGINNER'S BOX-2

1. 191.11 Cal. 2. 50 kcal 3. 1 cal
4. 10°C 5. 1 : 1 6. 615 kcal.
7. 4000 cal 8. 160°C
9. 7.5°C , water 10. (A)

BEGINNER'S BOX-3

- 1.(a) The thermal conductivity of brass is high, i.e., brass is a good conductor of heat. So, when a brass tumbler is touched, heat quickly flows from human body to tumbler. Consequently, the tumbler appears colder. On the other hand, wood is a bad conductor. So, heat does not flow from the human body to the wooden tray in this case. Thus, it appears comparatively hotter.
- (b) This is because the two thin cloth enclose a layer of air between them. Since air is a bad conductor of heat therefore the conduction of heat is prevented.
- (c) Because mudhouses are bad conductor of heat.
- (d) Because air is trapped between their wings since air is bad conductor of heat therefore the conduction of heat is prevented.
- (e) This is because wool contains air in its pores. Air is a bad conductor of heat. In fact, wool is also a bad conductor of heat so both the air and wool do not permit heat to be conducted away from. So woolen clothes are warm.
2. (a) 240 W (b) 1.5 3. 40°C
4. $\theta_1 = 116^\circ\text{C}$, $\theta_2 = 74^\circ\text{C}$
5. 68°C ; 0.9°C/cm 6. 7 : 2

BEGINNER'S BOX-4

- 1.(a) A body whose reflectivity is large would naturally absorb less heat. So, a body with large reflectivity is a poor absorber of heat. Poor absorbers are poor emitters. So, a body with large reflectivity is a poor emitter.
- (b) Let T be the temperature of the hot iron in the furnace. Heat radiated per second per unit area, $E = \sigma T^4$
When the body is placed in the open at temperature T_0 , then the heat radiated/second/area, $E' = \sigma (T^4 - T_0^4)$
Clearly $E' < E$. So, the optical pyrometer gives too low a value for the temperature in the open.
- (c) The lower layers of earth's atmosphere reflect infrared radiations from earth back to the surface of earth. Thus the heat radiation received by the earth from the sun during the day are kept trapped by the atmosphere. If atmosphere of earth were not there, its surface would become too cold to live.
- (d) When the electric heater is switched on, a stage is quickly reached when the rate at which heat is generated by electric. Current becomes equal to the rate at which heat is lost by radiation. This is case of thermal equilibrium.
- (e) Because black surface is good absorber of heat.
- (f) On a clear night, the earth radiates energy into space at a rate proportional to the fourth power of its temperature (about 300K). The incoming radiation from space is very small because its average temperature is nearly absolute zero. On the other hand with cloud over, the earth radiates at 300K, but the radiation is absorbed in the clouds, which radiate energy back to earth again the radiation is trapped, like the green house effect.
- (g) This is due to the fact that a vacuum is created between the two walls of the thermos flask. Heat can neither flow from inside the flask to outside nor from outside air to the liquid inside the flask. In this way, loss of heat by conduction and convection has been minimised. To minimise loss of heat by radiation the surface is made shining.

2. $\sqrt{\frac{r_2}{r_1}}$ 3. 40°C 4. 28°C
5. 1927 K

BEGINNER'S BOX-5

1. $\frac{1}{2}$ mole 2. 120 kPa 3. 2.2×10^{24}
4. (a) $\frac{N_1}{N_2} = \frac{3}{2}$; (b) 0.947

BEGINNER'S BOX-6

1. (a) 6.21×10^{-21} J; (b) 10.35×10^{-21} J; (c) 6231 J
2. (a) 127°C, (b) 527°C
3. (a) 3, (b) $\frac{5}{3}$, (c) 450 R
4. $v_{\text{avg}} = \frac{v_1 + v_2 + v_3}{3}$,
Root mean square speed = $\sqrt{\frac{v_1^2 + v_2^2 + v_3^2}{3}}$
5. 9.2 m/s
6. (a) $\frac{E_1}{E_2} = \frac{1}{1}$ ($\because T = \text{same for both}$);

$$(b) v_{\text{rms}} = \sqrt{\frac{70.9}{39.9}} = \frac{1.33}{1}$$

BEGINNER'S BOX-7

1. (a) $W_{\text{ABCA}} = \frac{1}{2} \times (2V_0 - V_0) \times (3P_0 - P_0) = P_0 V_0$
(b) $\frac{5}{2} P_0 V_0$ and $3P_0 V_0$, (c) $\frac{P_0 V_0}{2}$
2. -5×10^5 J

3. $W_{\text{AB}} = 6000$ J
 $W_{\text{BC}} = \text{zero}$
 $W_{\text{CD}} = 1000$ J
 $W_{\text{DA}} = \text{zero}$
4. (a) 24×10^2 J; (b) 36×10^2 J; (c) 60×10^2 J
5. 4.16 J/cal

BEGINNER'S BOX-8

1. 3650 J
2. (a) 162.5 cal; (b) 237.5 cal
3. (A) -p, (B) -q, (C) -r
4. (A) -q, (B) -r, (C) -p, (D) -s
5. (A) -p, s, (B) -p, q, r, s, (C) -s, (D) -q, s, (E) -r, s
6. (a) 700 J; (b) 500 J
7. (a) $T_B = 150$ K; $T_C = 37.5$ K; (b) 2400 J
8. 1152.6 joule.

BEGINNER'S BOX-9

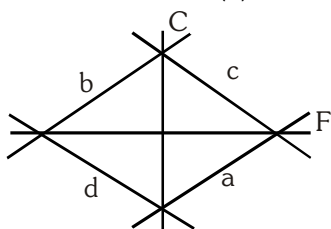
1. 168 J 2. 400 K
3. (a) 40%; (b) 240 K; (c) 400 J
4. 550 K 5. 250 K 6. 126 watt.

EXERCISE-I (Conceptual Questions)

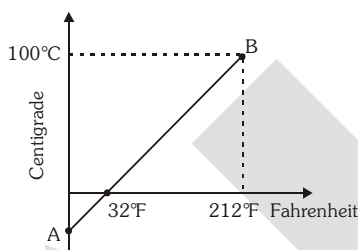
Build Up Your Understanding

TEMPERATURE & THERMAL EXPANSION

- At what temperature does the temperature in Celsius and Fahrenheit equalise
(1) -40° (2) 40° (3) 36.6° (4) 38°
- A difference of temperature of 25°C is equivalent to a difference of :
(1) 45°F (2) 72°F (3) 32°F (4) 25°F
- Which of the curves in figure represents the relation between Celsius and Fahrenheit temperature?
(1) Curve a (2) Curve b
(3) Curve c (4) Curve d



- The graph AB shown in figure is a plot of temperature of a body in degree Celsius and degree Fahrenheit. Then



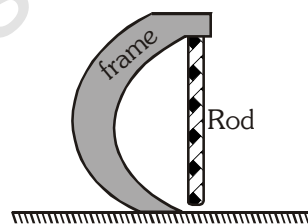
- slope of line AB is $9/5$
 - slope of line AB is $5/9$
 - slope of line AB is $1/9$
 - slope of line AB is $3/9$
- Oxygen boils at -183°C . This temperature is approximately in Fahrenheit is :-
(1) -329°F (2) -261°F
(3) -215°F (4) -297°F
- Using which of the following instrument, the temperature of the sun can be determined ?
(1) Platinum thermometer
(2) Gas thermometer
(3) Pyrometer
(4) Vapour pressure thermometer

- Two thermometers X and Y have ice points marked at 15° and 25° and steam points marked as 75° and 125° respectively. When thermometer X measures the temperature of a bath as 60° on it, what would thermometer Y read when it is used to measure the temperature of the same bath ?
(1) 60° (2) 75° (3) 100° (4) 90°
- The figure below shows four isotropic solids having positive coefficient of thermal expansion. A student predicts that on heating the solid following things can happen. Mark true (T) or False (F) for comments made by the student.



- The angle α in figure (1) will not change.
 - The length of line in figure (2) will decrease.
 - The radius of inner hole will decrease.
 - The distance AB will increase.
- T F F T
 - F T T F
 - T T T T
 - F F T F

- At STP a rod is hung from a frame as shown in figure, leaving a small gap between the rod and floor. The frame and rod system is heated uniformly upto 350 K . Then



- The rod will never touch the floor in any case.
 - If $\alpha_{\text{rod}} > \alpha_{\text{frame}}$, then rod may touch the floor.
 - If $\alpha_{\text{rod}} < \alpha_{\text{frame}}$, then rod may touch the floor.
 - None of the above
- A steel tape gives correct measurement at 20°C . A piece of wood is being measured with the steel tape at 0°C . The reading is 25 cm on the tape, the real length of the given piece of wood must be:
(1) 25 cm (2) $<25\text{ cm}$
(3) $>25\text{ cm}$ (4) can not say

11. The volume of a metal sphere increases by 0.15 % when its temperature is raised by 24°C . The coefficient of linear expansion of metal is :
 (1) $2.5 \times 10^{-5} / ^{\circ}\text{C}$ (2) $2.0 \times 10^{-5} / ^{\circ}\text{C}$
 (3) $-1.5 \times 10^{-5} / ^{\circ}\text{C}$ (4) $1.2 \times 10^{-5} / ^{\circ}\text{C}$
12. Suppose there is a hole in a copper plate. On heating the plate, diameter of hole, would :
 (1) always increase
 (2) always decrease
 (3) always remain the same
 (4) none of these
13. If two rods of length L and $2L$ having coefficients of linear expansion α and 2α respectively are connected so that total length becomes $3L$, the average coefficient of linear expansion of the composition rod equals:
 (1) $\frac{3}{2} \alpha$ (2) $\frac{5}{2} \alpha$
 (3) $\frac{5}{3} \alpha$ (4) none of these
14. The table gives the initial length ℓ_0 , change in temperature ΔT and change in length $\Delta \ell$ of four rods. Which rod has greatest coefficient of linear expansion
- | Rod | $\ell_0(\text{m})$ | $\Delta T(^{\circ}\text{C})$ | $\Delta \ell(\text{m})$ |
|-----|--------------------|------------------------------|-------------------------|
| A1 | 1 | 100 | 1 |
| A2 | 1 | 100 | 2 |
| A3 | 1.5 | 50 | 3 |
| A4 | 2.5 | 20 | 4 |
- (1) A_1 (2) A_2 (3) A_3 (4) A_4
15. An iron bar (Young's modulus = 10^{11} N/m^2 , $\alpha = 10^{-6} / ^{\circ}\text{C}$) 1 m long and 10^{-3} m^2 in area is heated from 0°C to 100°C without being allowed to bend or expand. Find the compressive force developed inside the bar.
 (1) 10,000 N (2) 1000 N
 (3) 5000 N (4) 10^5 N
16. A rod of length 2m rests on smooth horizontal floor. If the rod is heated from 0°C to 20°C . Find the longitudinal strain developed?
 ($\alpha = 5 \times 10^{-5} / ^{\circ}\text{C}$)
 (1) 10^{-3} (2) 2×10^{-3} (3) Zero (4) None

CALORIMETRY

17. A body of mass 5 kg falls from a height of 30 metre. If its all mechanical energy is changed into heat, then heat produced will be:-
 (1) 350 cal (2) 150 cal
 (3) 60 cal (4) 6 cal
18. A bullet moving with velocity v collides against wall. consequently half of its kinetic energy is converted into heat. If the whole heat is acquired by the bullet, the rise in temperature will be:-
 (1) $v^2/4S$ (2) $4v^2 / 2S$
 (3) $v^2 / 2S$ (4) v^2 / S
19. The amount of heat required in converting 1 g ice at -10°C into steam at 100°C will be :-
 (1) 3028 J (2) 6056 J
 (3) 721 J (4) 616 J
20. 2 kg ice at -20°C is mixed with 5 kg water at 20°C . Then final amount of water in the mixture would be;
 Given specific heat of ice = $0.5 \text{ cal/g}^{\circ}\text{C}$,
 Specific heat of water = $1 \text{ cal/g}^{\circ}\text{C}$,
 Latent heat of fusion for ice = 80 cal/g .
 (1) 6 kg (2) 5 kg (3) 4 kg (4) 2 kg
21. Two identical masses of 5 kg each fall on a wheel from a height of 10m. The wheel disturbs a mass of 2kg water, the rise in temperature of water will be :
 (1) 2.6°C (2) 1.2°C (3) 0.32°C (4) 0.12°C
22. A block of mass 2.5 kg is heated to temperature of 500°C and placed on a large ice block. What is the maximum amount of ice that can melt (approx.). Specific heat for the body = $0.1 \text{ cal/g}^{\circ}\text{C}$.
 (1) 1 kg (2) 1.5 kg (3) 2 kg (4) 2.5 kg
23. 1 kg of ice at -10°C is mixed with 4.4 kg of water at 30°C . The final temperature of mixture is : (specific heat of ice = $2100 \text{ J/kg}^{\circ}\text{C}$)
 (1) 2.3°C (2) 4.4°C (3) 5.3°C (4) 8.7°C
24. The amount of heat required to convert 1 g of ice at 0°C into steam at 100°C , is
 (1) 716 cal. (2) 500 cal.
 (3) 180 cal. (4) 100 cal.

25. The latent heat for vapourisation for 1 g water is 536 cal. Its value in Joule/kg will be :-

- (1) 2.25×10^6 (2) 2.25×10^3
(3) 2.25 (4) None of these

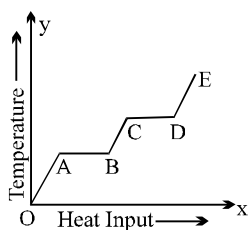
26. If 10 g ice at 0°C is mixed with 10 g water at 20°C , the final temperature will be :-

- (1) 50°C (2) 10°C (3) 0°C (4) 15°C

27. 420 joule of energy supplied to 10 g of water will raise its temperature by nearly :-

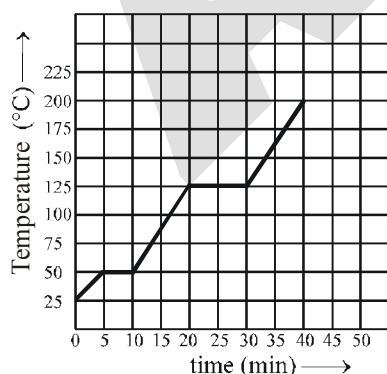
- (1) 1°C (2) 4.2°C (3) 10°C (4) 42°C

28. A solid material is supplied with heat at a constant rate. The temperature of material is changing with heat input as shown in the figure. What does slope DE represent.



- (1) latent heat of liquid
(2) latent heat of vapour
(3) heat capacity of vapour
(4) inverse of heat capacity of vapour

29. The graph shown in the figure represent change in the temperature of 5 kg of a substance as it absorbs heat at a constant rate of 42 kJ min^{-1} . The latent heat of vapourization of the substance is :



- (1) 630 kJ kg^{-1} (2) 126 kJ kg^{-1}
(3) 84 kJ kg^{-1} (4) 12.6 kJ kg^{-1}

30. A block of ice with mass m falls into a lake. After impact, a mass of ice $m/5$ melts. Both the block of ice and the lake have a temperature of 0°C . If L represents the heat of fusion, the minimum distance the ice fell before striking the surface is

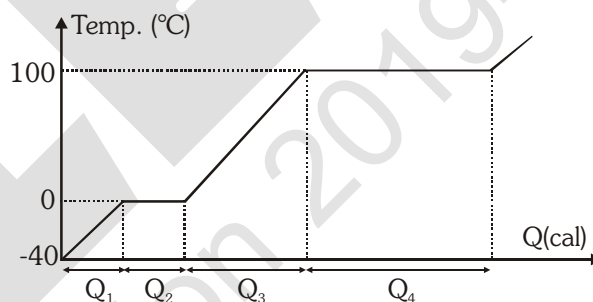
- (1) $\frac{L}{5g}$ (2) $\frac{5L}{g}$ (3) $\frac{gL}{5m}$ (4) $\frac{mL}{5g}$

31. 10 g of ice at 0°C is kept in a calorimeter of water equivalent 10 g. How much heat should be supplied to the apparatus to evaporate the water thus formed? (Neglect loss of heat)

- (1) 6200 cal (2) 7200 cal
(3) 13600 cal (4) 8200 cal

32. Figure shows the temperature variation when heat is added continuously to a specimen of ice (10 g) at -40°C at constant rate.

(Specific heat of ice = $0.53 \text{ cal/g } ^\circ\text{C}$ and $L_{\text{ice}} = 80 \text{ cal/g}$, $L_{\text{water}} = 540 \text{ cal/g}$)



Column-I

Column-II

- (A) Value of Q_1 (in cal) (P) 800
(B) Value of Q_2 (in cal) (Q) 1000
(C) Value of Q_3 (in cal) (R) 5400
(D) Value of Q_4 (in cal) (S) 212
(T) 900

- (1) $A \rightarrow S$; $B \rightarrow P$; $C \rightarrow Q$; $D \rightarrow T$
(2) $A \rightarrow P$; $B \rightarrow S$; $C \rightarrow Q$; $D \rightarrow R$
(3) $A \rightarrow P$; $B \rightarrow S$; $C \rightarrow R$; $D \rightarrow Q$
(4) $A \rightarrow S$; $B \rightarrow P$; $C \rightarrow Q$; $D \rightarrow R$

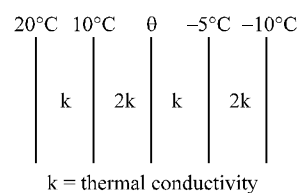
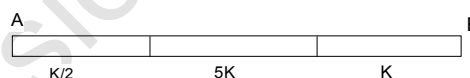
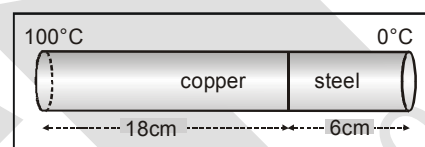
33. The thermal capacity of any body is

- (1) a measure of its capacity to absorb heat
(2) a measure of its capacity to provide heat
(3) the quantity of heat required to raise its temperature by a unit degree
(4) the quantity of heat required to raise the temperature of a unit mass of the body by a unit degree

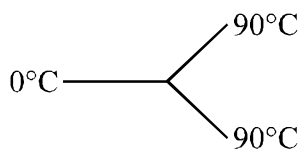
34. 2 litre water at 27°C is heated by a 1 kW heater in an open container. On an average heat is lost to surroundings at the rate 160 J/s. The time required for the temperature to reach 77°C is
 (1) 8 min 20 sec (2) 10 min
 (3) 7 min (4) 14 min
35. A piece of ice (heat capacity = $2100 \text{ J kg}^{-1} ^\circ\text{C}^{-1}$ and latent heat = $3.36 \times 10^5 \text{ J kg}^{-1}$) of mass m grams is at -5°C at atmospheric pressure. It is given 420 J of heat so that the ice starts melting. Finally when the ice-water mixture is in equilibrium, it is found that 1 g of ice has melted. Assuming there is no other heat exchange in the process, the value of m is
 (1) 40 g (2) 8 g (3) 16 g (4) 24 g
36. A 2100 W continuous flow geyser (instant geyser) has water inlet temperature = 10°C while the water flows out at the rate of 20 g/s. The outlet temperature of water must be about
 (1) 20°C (2) 30°C (3) 35°C (4) 40°C
37. A continuous flow water heater (geyser) has an electrical power rating = 2 kW and efficiency of conversion of electrical power into heat = 80%. If water is flowing through the device at the rate of 100 cc/s, and the inlet temperature is 10°C , the outlet temperature will be
 (1) 12.2°C (2) 13.8°C
 (3) 20°C (4) 16.5°C

CONDUCTION AND CONVECTION

38. The ratio of coefficient of thermal conductivity of two different materials is 5:3. If the thermal resistance of rods of same area of these material is same, then what is ratio of length of these rods -
 (1) 3:5 (2) 5:3 (3) 25:9 (4) 9:25
39. Rate of heat flow through a cylindrical rod is Q_1 . Temperatures of ends of rod are T_1 and T_2 . If all the linear dimensions of the rod become double and temperature difference remains same, its rate of heat flow is Q_2 , then :-
 (1) $Q_1 = 2Q_2$ (2) $Q_2 = 2Q_1$
 (3) $Q_2 = 4Q_1$ (4) $Q_1 = 4Q_2$
40. A heat flux of 4000 J/s is to be passed through a copper rod of length 10 cm and area of cross section 100 cm^2 . The thermal conductivity of copper is $400 \text{ W/m}^\circ\text{C}$. The two ends of this rod must be kept at a temperature difference of
 (1) 1°C (2) 10°C
 (3) 100°C (4) 1000°C
41. The coefficient of thermal conductivity of copper is nine times that of steel. In the composite cylindrical bar shown in the figure what will be the temperature at the junction of copper and steel ?
 (1) 75°C
 (2) 67°C
 (3) 33°C
 (4) 25°C
42. A composite rod made of three rods of equal length and cross-section as shown in the fig. The thermal conductivities of the materials of the rods are $K/2$, $5K$ and K respectively. The end A and end B are at constant temperatures. All heat entering the face A goes out of the end B and there being no loss of heat from the sides of the bar. The effective thermal conductivity of the bar is
 (1) $15K/16$ (2) $6K/13$
 (3) $5K/16$ (4) $2K/13$
43. The figure shows the face and interface temperature of a composite slab containing of four layers of two materials having identical thickness. Under steady state condition, find the value of temperature θ .

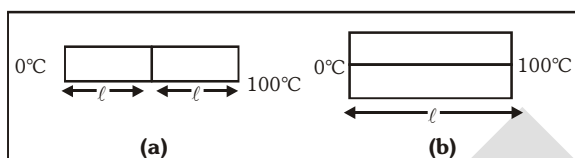


44. Three rods made of the same material and having the same cross-section have been joined as shown in the figure. Each rod is of the same length. The left and right ends are kept at 0°C and 90°C respectively. The temperature of the junction of the three rods will be :



- (1) 45°C (2) 60°C (3) 30°C (4) 20°C

45. Two identical square rods of metal are welded end to end as shown in figure (1), 20 calories of heat flows through it in 4 minutes. If the rods are welded as shown in figure (2), the same amount of heat will flow through the rods in -



- (1) 1 minute (2) 2 minutes
(3) 4 minutes (4) 16 minutes

46. The coefficient of thermal conductivity depends upon-
- (1) Temperature difference of two ends
 - (2) Area of the plate
 - (3) Thickness of the plate
 - (4) Material of the plate
47. Which of the following cylindrical rods will conduct most heat, when their ends are maintained at the same steady temperature
- (1) Length 1 m; radius 1 cm
 - (2) Length 2 m; radius 1 cm
 - (3) Length 2 m; radius 2 cm
 - (4) Length 1 m; radius 2 cm

48. Gravitational force is required for -
- (1) Stirring of liquid (2) Convection
 - (3) Conduction (4) Radiation

49. The layers of atmosphere are heated through -
- (1) Convection (2) Conduction
 - (3) Radiation (4) 2 and 3 both

50. The lengths and radii of two rods made of same material are in the ratios 1 : 2 and 2 : 3 respectively. If the temperature difference between the ends for the two rods be the same then in the steady state. The amount of heat flowing per second through them will be in the ratio of

- (1) 1 : 3 (2) 4 : 3
(3) 8 : 9 (4) 3 : 2

51. Two metal rods, 1 & 2 of same length have same temperature difference between their ends, their thermal conductivities are K_1 & K_2 and cross sectional areas A_1 & A_2 respectively. What is required condition for same rate of heat conduction in them.

- (1) $K_1 = K_2$ (2) $K_1 A_1 = K_2 A_2$
(3) $\frac{K_1}{A_1} = \frac{K_2}{A_2}$ (4) $\frac{K_1}{\ell_1^2} = \frac{K_2}{\ell_2^2}$

52. The temperature of hot and cold end of a 20 cm long rod in thermal steady state are at 100°C and 20°C respectively. Temperature at the centre of the rod is

- (1) 50°C (2) 60°C
(3) 40°C (4) 30°C

53. Consider a compound slab consisting of two different materials having equal thicknesses and thermal conductivities K and $2K$, respectively. The equivalent thermal conductivity of the slab is

- (1) $3K$ (2) $\frac{4}{3}K$
(3) $\frac{2}{3}K$ (4) $\sqrt{2}K$

54. Under steady state, the temperature of a body

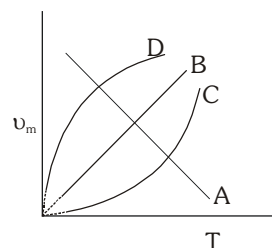
- (1) Increases with time
- (2) Decreases with time
- (3) Does not change with time and is same at all the points of the body
- (4) Does not change with time but is different at different points of the body

67. It is hotter at the same distance over the top of a fire than it is in the side of it, mainly because
 (1) Air conducts heat upwards
 (2) Heat is radiated upwards
 (3) Convection takes more heat upwards
 (4) Convection, conduction and radiation all contribute significantly transferring heat upward

RADIATION

68. A spherical body of area A , and emissivity $e = 0.6$ is kept inside a black body. What is the rate at which energy is radiated per second at temperature T .
 (1) $0.6 \sigma AT^4$ (2) $0.4 \sigma AT^4$
 (3) $0.8 \sigma AT^4$ (4) $1.0 \sigma AT^4$
69. Radius of two spheres of same material are 1 & 4 m respectively and their temperature are 4×10^3 and 2×10^3 K respectively. Then ratio of emitted energy of spheres per sec. will be -
 (1) 1:2 (2) 2:1 (3) 1:1 (4) 4:1
70. Cooling rate of a sphere of 600 K at external environment (200 K) is R . When the temperature of sphere is reduced to 400 K then cooling rate of the sphere becomes :
 (1) $\frac{3}{16} R$ (2) $\frac{16}{3} R$ (3) $\frac{9}{27} R$ (4) None
71. Temperature of a body is 400°C . Assuming the surrounding temperature to be negligible. At what temperature will body emit double energy radiation?
 (1) 200°C (2) 200 K
 (3) 800°C (4) 800 K
72. If temperature of ideal black body increased by 10%, then percentage increase in quantity of radiation emitted from its surface will be :-
 (1) 10% (2) 40% (3) 46% (4) 100%
73. The rectangular surface of area $8\text{cm} \times 4\text{cm}$ of a black body at a temperature of 127°C emits energy at the rate of E . If the length and breadth of the surface are each reduced to half of the initial value and the temperature is raised to 327°C , the rate of emission of energy will become.
 (1) $\frac{3}{8}E$ (2) $\frac{81}{16}E$ (3) $\frac{9}{16}E$ (4) $\frac{81}{64}E$

74. The rate of emission of radiation of a black body at 273°C is E , then the rate of emission of radiation of this body at 0°C will be
 (1) $\frac{E}{16}$ (2) $\frac{E}{4}$ (3) $\frac{E}{8}$ (4) 0
75. If a liquid takes 30 s in cooling from 95°C to 90°C and 70 s in cooling from 55°C to 50°C then temperature of room is -
 (1) 16.5°C (2) 22.5°C
 (3) 28.5°C (4) 32.5°C
76. The thermal capacities of two bodies are in the ratio of 1:4. If the rate of loss of heat are equal for the two bodies under identical conditions of surroundings, then the ratio of rate of fall of temperature of the two bodies is -
 (1) 1:4 (2) 4:1
 (3) 1:8 (4) 8:1
77. Newton's law of cooling is used in laboratory for the determination of the
 (1) Specific heat of the gases
 (2) The latent heat of gases
 (3) Specific heat of liquids
 (4) Latent heat of liquids
78. The $\nu_m - T$ curve for a perfect black body is -
 ($\nu_m \rightarrow$ frequency corresponding to maximum emission of radiation)



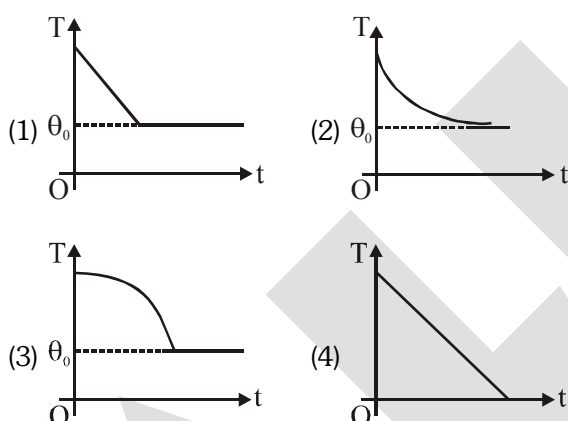
- (1) A (2) B (3) C (4) D
79. Two stars appear to be red and blue, what is true about them -
 (1) The red star is nearer
 (2) The blue star is nearer
 (3) The temperature of red star is more
 (4) The temperature of blue star is more

80. The temperature of a furnace is 2324°C and the intensity is maximum in its radiation spectrum nearly at 12000 \AA . If the intensity in the spectrum of a star is maximum nearly at 4800 \AA , then the surface temperature of star is
- (1) 8400°C (2) 7200°C
(3) 6219.5°C (4) 5900°C
81. There is a black spot on a body. If the body is heated and carried in a dark room then it glows more. This can be explained on the basis of –
- (1) Newton's law of cooling
(2) Wein's law
(3) Kirchhoff's law
(4) Stefan's law
82. The colour of a star is an indication of its –
- (1) Weight (2) Distance
(3) Temperature (4) Size
83. If a carved black utensil is heated to high temperature and then brought in dark then :
- (1) Both utensil and its carving will shine
(2) Only carving will shine
(3) Only utensil will shine
(4) None of the utensil and carving will shine
84. According to Newton's law of cooling, the rate of cooling of a body is proportional to :–
- (1) Temperature of the body
(2) Temperature of the surrounding
(3) Fourth power of the temperature of body
(4) Difference of the temperature of the body and the surrounding.
85. The original temperature of a black body is 727°C . Calculate temperature at which total radiant energy from this black body becomes double :
- (1) 971 K (2) 1190 K
(3) 2001 K (4) 1458 K
86. Calculate the energy radiated per minute from the filament of an incandescent lamp at 2000K if the surface area is $5 \times 10^{-5} \text{ m}^2$ and its relative emittance is 0.85, $\sigma = 5.7 \times 10^{-8} \text{ MKS units}$:
- (1) 1230 J (2) 2215 J
(3) 2115 J (4) 2325 J
87. Ratio of radius of curvature of cylindrical emitters of same material is 1:4 and their temperature are in ratio 2:1. Then ratio of amount of heat emitted by them is - (For Cylinder length = radius)
- (1) 2:1 (2) 1:1
(3) 4:1 (4) 1:4
88. The ideal black body is :
- (1) Hot coal at high temperature
(2) Surface of glass printed with coaltar
(3) Metal surface
(4) A hollow container painted with black colour
89. The energy emitted per second by a black body at 27°C is 10 J. If temperature of the black body is increased to 327°C , the energy emitted per second will be :-
- (1) 20 J (2) 40 J
(3) 80 J (4) 160 J
90. Energy is being emitted from the surface of black body at 127°C at the rate of $1.0 \times 10^6 \text{ J/s m}^2$. The temperature of black body at which the rate of energy is $16.0 \times 10^6 \text{ J/s m}^2$ will be :
- (1) 754°C (2) 527°C
(3) 254°C (4) 508°C
91. Solar constant for earth is 2 cal/min cm^2 , if distance of mercury from sun is 0.4 times than distance of earth from sun then solar constant for mercury will be?
- (1) $12.5 \text{ cal/min cm}^2$
(2) 25 cal/min cm^2
(3) $0.32 \text{ cal/min cm}^2$
(4) 2 cal/min cm^2

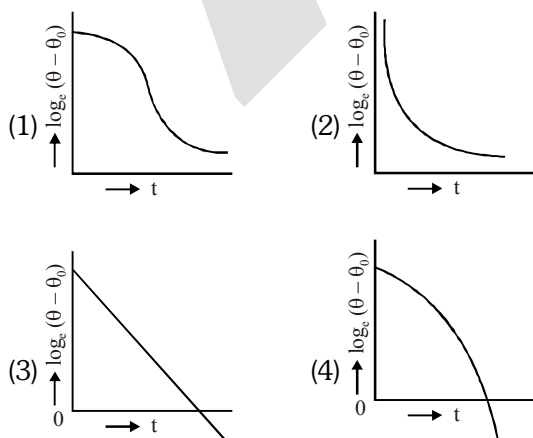
- 92.** Two spherical bodies A (radius 6 cm) and B (radius 18 cm) are at temperature T_1 and T_2 respectively. The maximum intensity in the emission spectrum of A is at 500 nm and in that of B is at 1500 nm. Considering them to be black bodies, what will be the ratio of the rate of total energy radiated by A to that of B ?
 (1) 9 (2) 6 (3) 12 (4) 3

- 93.** Star S_1 emits maximum radiation of wavelength 420 nm and the star S_2 emits maximum radiation of wavelength 560 nm, what is the ratio of the temperature of S_1 and S_2 :
 (1) $4/3$ (2) $(4/3)^{1/4}$
 (3) $3/4$ (4) $(3/4)^{1/2}$

- 94.** If a piece of metal is heated to temperature θ and then allowed to cool in a room which is at temperature θ_0 , the graph between the temperature T of the metal and time t will be closest to



- 95.** A liquid in a beaker has temperature θ at time t and θ_0 is temperature of surroundings, then according to Newton's law of cooling, correct graph between $\log_e(\theta - \theta_0)$ and t is:



- 96.** A bucket full of hot water cools from 75°C to 70°C in time T_1 , from 70° to 65°C in time T_2 and from 65°C to 60°C in time T_3 , then
 (1) $T_1 = T_2 = T_3$ (2) $T_1 > T_2 > T_3$
 (3) $T_1 < T_2 < T_3$ (4) $T_1 > T_2 < T_3$

- 97.** The Wein's displacement law express relation between :-
 (1) Wavelength corresponding to maximum energy and temperature.
 (2) Radiation energy and wavelength
 (3) Temperature and wavelength
 (4) Colour of light and temperature

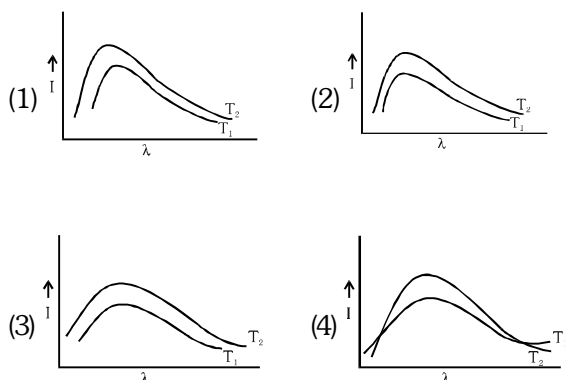
- 98.** Four identical calorimeters painted in different colours, are heated to same temperature and then allowed to cool in vacuum. Which will cool fastest ?
 (1) One which is painted bright
 (2) One which is painted thick white
 (3) One which is painted thick black
 (4) One which is painted bright white

- 99.** A body cools from 60°C to 50°C in 10 minutes. If the room temperature is 25°C and assuming Newton's cooling law holds good, the temperature of the body at the end of next 10 minutes is :
 (1) 45°C (2) 42.85°C
 (3) 40°C (4) 38.5°C

- 100.** As compared to the person with white skin, the person with black skin will experience
 (1) Less heat and more cold
 (2) More heat and more cold
 (3) More heat and less cold
 (4) Less heat and less cold

- 101.** We consider the radiation emitted by the human body. Which of the following statements is true ?
 (1) The radiation is emitted during the summers and absorbed during the winters
 (2) The radiation emitted lies in the ultraviolet region and hence is not visible
 (3) The radiation emitted is in the infra-red region
 (4) The radiation is emitted only during the day

- 102.** Shown below are the black body radiation curves at temperatures T_1 and T_2 ($T_2 > T_1$). Which of the following plots is correct :-



- 103.** The radii of two spheres made of same metal are r and $2r$. These are heated to the same temperature and placed in the same surrounding. The ratio of rates of decrease of their temperature will be
(1) 1 : 1 (2) 4 : 1 (3) 1 : 4 (4) 2 : 1

- 104.** If E is the total energy emitted by a body at a temperature T K and $E_{\lambda, \max}$ is the maximum spectral energy emitted by it at the same temperature, then-
(1) $E \propto T^4$; $E_{\lambda, \max} \propto T^5$ (2) $E \propto T^4$; $E_{\lambda, \max} \propto T^{-5}$
(3) $E \propto T^{-4}$; $E_{\lambda, \max} \propto T^4$ (4) $E \propto T^5$; $E_{\lambda, \max} \propto T^4$

- 105.** If e_λ and a_λ be the emissive power and absorption power respectively of a body and E_λ be the emissive power of an ideal black body, then from Kirchhoff's laws
(1) $a_\lambda = E_\lambda / e_\lambda$ (2) $a_\lambda / e_\lambda = E_\lambda$
(3) $e_\lambda / a_\lambda = E_\lambda$ (4) $e_\lambda = E_\lambda / a_\lambda$

- 106.** A liquid takes 5 min. to cool from 80°C to 50°C . How much time it will take to cool from 60°C to 30°C . Temperature of surroundings is 20°C -
(1) 15 min. (2) 20 min. (3) 100 min. (4) 9 min.

- 107.** A cup of tea cools from 80°C to 60°C in one minute. The ambient temperature is 30°C . In cooling from 60°C to 50°C , it will take :
(1) 50 sec. (2) 90 sec (3) 60 sec. (4) 48 sec.

- 108.** If λ_m denotes the wavelength at which the radiation emission from a black body at a temperature T K is maximum then :
(1) $\lambda_m \propto T$ (2) $\lambda_m \propto T^2$
(3) $\lambda_m \propto T^{-1}$ (4) $\lambda_m \propto T^{-2}$

- 109.** A body is in thermal equilibrium with the surrounding :
(1) It will stop emitting heat radiation
(2) Amount of radiation emitted and absorbed by it will be equal
(3) It will emit heat radiation at faster rate
(4) It will emit heat radiation slowly

- 110.** Which of the following statement is correct for ideal black body :
(1) This absorbs visible radiation only.
(2) This absorbs infrared radiation only
(3) This absorbs half of radiation only and reflects the half
(4) This totally absorbs heat radiation of all the wavelengths

- 111.** Two spheres P and Q of same colour having radii 8 cm and 2 cm are maintained at temperatures 127°C and 527°C respectively. The ratio of energy radiated by P and Q is -
(1) 0.054 (2) 0.0034
(3) 1 (4) 2

- 112.** On increasing the temperature of a black body, wavelength for maximum emission.
(1) Shifts towards smaller wavelength
(2) Shifts towards greater wavelength
(3) Does not shift
(4) Depends on the shape of source.

- 113.** A solid cube and sphere are made of same substance and both have same surface area. If the temperature of both bodies 120°C then :
(1) Both will loss of Heat by same rate
(2) Rate of loss of Heat of cube will be more than that of the sphere
(3) Rate of loss of Heat of the sphere will be more than that of the cube
(4) Rate of loss of Heat will be more for that whose mass is more

- 114.** Two spheres of radii in the ratio 1 : 2 and densities in the ratio 2 : 1 and of same specific heat, are heated to same temperature and left in the same surrounding. Their rate of falling temperature will be in the ratio :
(1) 2 : 1 (2) 1 : 1
(3) 1 : 2 (4) 1 : 4

115. A black body at 200 K is found to emit maximum energy at a wavelength $14 \mu\text{m}$. When its temperature is raised to 1000 K, then wavelength at which maximum energy emitted is :

- (1) $14 \mu\text{m}$ (2) $15 \mu\text{m}$
(3) $2.8 \mu\text{m}$ (4) $28 \mu\text{m}$

116. The spectrum from a black body radiation is a :

- (1) line spectrum
(2) band spectrum
(3) continuous spectrum
(4) line & band both

117. The temperature of a perfect black body is 727°C and its area is 0.1 m^2 . If stefan's constant is $5.67 \times 10^{-8} \text{ watt/m}^2\text{-K}^4$, then heat radiated by it in 1 minute is :

- (1) 8100 cal (2) 81000 cal
(3) 810 cal (4) 81 cal

118. A black body radiates energy at the rate of $E \text{ watt/m}^2$ at a high temperature $T \text{ K}$. When the

temperature is reduced to $\frac{T}{2} \text{ K}$, the radiant energy will be

- (1) $\frac{E}{16}$ (2) $\frac{E}{4}$
(3) $4 E$ (4) $16 E$

119. In solar spectrum fraunhofer's lines are presents because :

- (1) Definite absorption takes place in photosphere of sun.
(2) Definite absorption takes place in chromosphere of sun.
(3) These wave lengths are not at all emitted by sun.
(4) Nuclear reactions take place in sun.

120. On reducing temperature of surface to one third, amount of radiation becomes :-

- (1) $\frac{1}{27}$ (2) $\frac{1}{81}$ (3) $\frac{1}{9}$ (4) $\frac{1}{3}$

121. The absorptive power of a perfectly black body is

- (1) zero (2) infinity
(3) 1.5 (4) 1.0

KTG & GAS LAWS AND IDEAL GAS EQUATION

122. Find the approximate number of molecules contained in a vessel of volume 7 litres at 0°C at $1.3 \times 10^5 \text{ pascals}$

- (1) 2.4×10^{23} (2) 3×10^{23}
(3) 6×10^{23} (4) 4.8×10^{23}

123. A real gas behaves like an ideal gas if its

- (1) pressure and temperature are both high
(2) pressure and temperature are both low
(3) pressure is high and temperature is low
(4) pressure is low and temperature is high

124. The lowest pressure (the best vacuum) that can be created in laboratory at 27°C is $10^{-11} \text{ mm of Hg}$. At this pressure, the number of ideal gas molecules per cm^3 will be :-

- (1) 3.22×10^{12} (2) 1.61×10^{12}
(3) 3.21×10^6 (4) 3.22×10^5

125. Two gases of equal molar amount are in thermal equilibrium. If P_a, P_b and V_a, V_b are their respective pressures and volumes, then which relation is true:-

- (1) $P_a \neq P_b, V_a = V_b$
(2) $V_a = V_b, V_a \neq V_b$
(3) $P_a/V_b = P_b/V_a$
(4) $P_a V_a = P_b V_b$

126. Equal volume of H_2, O_2 and He gases are at same temperature and pressure. Which of these will have large number of molecules :-

- (1) H_2
(2) O_2
(3) He
(4) All the gases will have same number of molecules

127. Gases obey vander - waal's equation at :

- (1) Only normal temperature and pressure
(2) Only high temperature and high pressure
(3) Only high temperature and low pressure
(4) All temperature and pressure

128. A box contains N molecules of a gas. If the number of molecules is doubled, then the pressure will :-

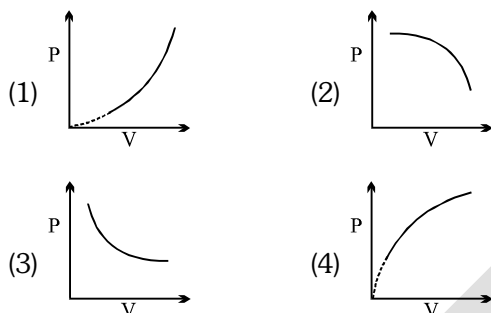
- (1) Decrease (2) Be same
(3) Be doubled (4) Get tripled

- 129.** An ideal gas mixture filled inside a balloon expands according to the relation $PV^{2/3} = \text{constant}$. The temperature inside the balloon is
 (1) increasing (2) decreasing
 (3) constant (4) can't be said

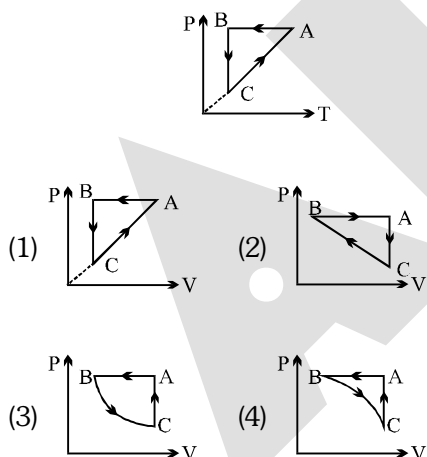
- 130.** A rigid tank contains 35 kg of nitrogen at 6 atm. Sufficient quantity of oxygen is supplied to increase the pressure to 9 atm, while the temperature remains constant. Amount of oxygen supplied to the tank is :

- (1) 5 kg (2) 10 kg
 (3) 20 kg (4) 40 kg

- 131.** An ideal gas follows a process $PT = \text{constant}$. The correct graph between pressure & volume is :-



- 132.** A cyclic process ABCA is shown in P-T diagram. When presented on P-V, it would



- 133.** 28 g of N_2 gas is contained in a flask at a pressure 10 atm. and at a temperature of 57°C . It is found that due to leakage in the flask, the pressure is reduced to half and the temperature reduced to 27°C . The quantity of N_2 gas that leaked out is -

- (1) 11/20 g (2) 80/11 g
 (3) 5/63 g (4) 63/5 g

- 134.** When temperature of a gas, contained in closed vessel is increased by 5°C , its pressure increases by 1%. The original temperature of the gas was approximately :

- (1) 500°C (2) 273°C
 (3) 227°C (4) 50°C

- 135.** During an experiment an ideal gas obeys an addition equation of state $P^2V = \text{constant}$. The initial temperature and volume of gas are T and V respectively. When it expands to volume 2V, then its temperature will be :

- (1) T (2) $\sqrt{2} T$
 (3) 2 T (4) $2\sqrt{2} T$

- 136.** 250 litre of an ideal gas is heated at constant pressure from 27°C such that its volume becomes 500 liters. The final temperature is :

- (1) 54°C (2) 300°C
 (3) 327°C (4) 600°C

- 137.** A balloon contains 500 m^3 of helium at 27°C and 1 atmosphere pressure. The volume of the helium at -3°C temperature and 0.5 atmosphere pressure will be-

- (1) 500 m^3 (2) 700 m^3
 (3) 900 m^3 (4) 1000 m^3

- 138.** A vessel has 6g of oxygen at pressure P and temperature 400 K. A small hole is made in it so that oxygen leaks out. How much oxygen leaks out if the final pressure is $P/2$ and temperature is 300 K ?

- (1) 3g (2) 2g (3) 4g (4) 5g

- 139.** Relation $PV = RT$ is given for following condition for real gas -

- (1) High temperature and high density
 (2) Low temperature and low density
 (3) High temperature and low density
 (4) Low temperature and high density

- 140.** A container of 5 litre has a gas at pressure of 0.8 m column of Hg. This is joined to an evacuated container of 3 litre capacity. The resulting pressure will be :- (At constant temp.)

- (1) 4/3 (2) 0.5 m
 (3) 2.0 m (4) 3/4 m

141. At a given temperature, the pressure of an ideal gas of density ρ is proportional to -

- (1) $\frac{1}{\rho^2}$ (2) $\frac{1}{\rho}$ (3) ρ^2 (4) ρ

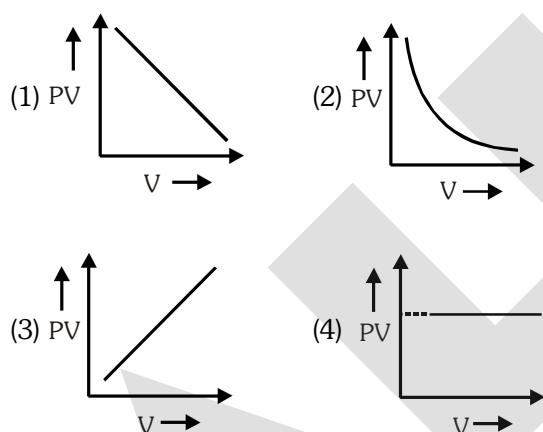
142. O_2 gas is filled in a cylinder. When pressure is increased 2 times, temperature becomes four times. then how much times their density will become :

- (1) 2 (2) 4 (3) $\frac{1}{4}$ (4) $\frac{1}{2}$

143. On increasing the temperature of a gas filled in a closed container by $1^\circ C$ its pressure increases by 0.4%, then initial temperature of the gas is-

- (1) $25^\circ C$ (2) $250^\circ C$
(3) 250 K (4) $2500^\circ C$

144. The variation of PV graph with V of a fixed mass of an ideal gas at constant temperature is graphically represented as shown in figure :



145. The number of oxygen molecules in a cylinder of volume 1 m^3 at a temperature of $27^\circ C$ and pressure 13.8 Pa is :

(Boltzmann's constant $k = 1.38 \times 10^{-23} \text{ JK}^{-1}$)

- (1) 6.23×10^{26} (2) 0.33×10^{28}
(3) 3.3×10^{21} (4) None of these

146. A cylinder contains 10 kg of gas at pressure of 10^7 N/m^2 . When final pressure is reduced to $2.5 \times 10^6 \text{ N/m}^2$ then quantity of gas taken out of the cylinder will be : (temperature of gas is constant)

- (1) 15.2 kg (2) 3.7 kg
(3) zero (4) 7.5 kg

147. Hydrogen and helium gases of volume V at same temperature T and same pressure P are mixed to have same volume V. The resulting pressure of the mixture will be :

- (1) $P/2$
(2) P
(3) 2P
(4) Depending on the relative mass of the gases

148. The equation of state for 5g of oxygen at a pressure P and temperature T occupying a volume V, will be :- (where R is the gas constant)

- (1) $PV = 5 RT$
(2) $PV = (5/2) RT$
(3) $PV = (5/16) RT$
(4) $PV = (5/32) RT$

149. In kinetic theory of gases, it is assumed that molecules :-

- (1) Have same mass but can have different volume
(2) Have same volume but masses can be different
(3) Have both mass and volume different
(4) Have same mass but negligible volume

150. The volume of an ideal gas is V at pressure P and temperature T. The mass of each molecule of the gas is m. The density of gas will be :- (K is Boltzmann's constant)

- (1) mKT (2) Pm / KT
(3) P / KTV (4) P / KT

151. The thermodynamic variables of a jar filled with gas A are P, V and T and another jar B filled with another gas are 2P, V/4 and 2T, where the symbols have their usual meaning. The ratio of the number of molecules of jar A to those of jar B is :

- (1) 4 : 1 (2) 2 : 1 (3) 1 : 2 (4) 1 : 1

152. At N.T.P. volume of a gas is changed to one fourth volume, at constant temperature then the new pressure will be :

- (1) 2 atm. (2) $2^{5/3}$ atm.
(3) 4 atm. (4) 1 atm.

153. A gas contained in a box of 0.1 m^3 at atmospheric pressure is connected to another vessel of 0.09 m^3 . Consequent change in pressure is X mm of Hg. Then X in metre is -

- (1) 0.4 (2) 0.5 (3) 0.36 (4) 0.3

- 167.** The reason for the absence of atmosphere on moon is that the :
- Value of v_{rms} of the molecules of gas is more than the value of escape velocity
 - Value of v_{rms} of gas is less than escape velocity
 - Value of v_{rms} is negligible
 - None of the above
- 168.** The speeds of 5 molecules of a gas (in arbitrary units) are as follows 2,3,4,5,6 The root mean square speed for these molecules is -
- 2.91
 - 3.52
 - 4.00
 - 4.24
- 169.** The root mean square speed of the molecules of a gas is :
- Independent of its pressure but directly proportional to its Kelvin temperature
 - Directly proportional to the square roots of both its pressure and its Kelvin temperature
 - Independent of its pressure but directly proportional to the square root of its Kelvin temperature
 - Directly proportional to both its pressure and its Kelvin temperature
- 170.** At 0°C temperature root mean square speed of which of the following gases will be maximum:-
- H_2
 - N_2
 - O_2
 - SO_2
- 171.** The root mean square speed of hydrogen molecules of an ideal hydrogen gas kept in a gas chamber is 3180 m/s . The pressure of the hydrogen gas is :- (Density of hydrogen gas = $8.99 \times 10^{-2} \text{ Kg/m}^3$, 1 atmosphere = $1.01 \times 10^5 \text{ N/m}^2$)
- 1.0 atmosphere
 - 1.5 atmosphere
 - 2.0 atmosphere
 - 3.0 atmosphere
- 172.** If the rms velocity of molecules of a gas in a container is doubled then the pressure will:-
- Become four times
 - Also get doubled
 - Be same
 - Become one half
- 173.** The root mean square velocity of a gas molecule of mass m at a given temperature is proportional to -
- m^0
 - m
 - \sqrt{m}
 - $\frac{1}{\sqrt{m}}$
- 174.** v_{rms} , v_{av} and v_{mp} are root mean square, average and most probable speeds of molecules of a gas obeying Maxwell's velocity distribution. Which of the following statements is correct
- $v_{rms} < v_{av} < v_{mp}$
 - $v_{rms} > v_{av} > v_{mp}$
 - $v_{mp} < v_{rms} < v_{av}$
 - $v_{mp} > v_{rms} > v_{av}$
- 175.** If the r.m.s. velocity of hydrogen becomes equal to the escape velocity from the earth surface, then the temperature of hydrogen gas would be-
- 1060 K
 - 5030 K
 - 8270 K
 - 10^4 K
- 176.** The pressure exerted by a gas in P_0 . If the mass of molecules becomes half and their velocities become double, then pressure will become
- $\frac{P_0}{2}$
 - P_0
 - $2P_0$
 - $4P_0$
- 177.** The root mean square (rms) speed of oxygen molecules O_2 at a certain temperature T (absolute) is v . If the temperature is doubled and oxygen gas dissociates into atomic oxygen. The rms speed :
- becomes $v/\sqrt{2}$
 - remains v
 - becomes $\sqrt{2}v$
 - becomes $2v$
- 178.** If the root mean square speed of hydrogen molecules is equal to root mean square speed of oxygen molecules at 47°C , the temperature of hydrogen is -
- 20 K
 - 47 K
 - 50 K
 - 94 K
- 179.** The root mean square and most probable speed of the molecules in a gas are :
- same
 - different
 - cannot say
 - depends on nature of the gas
- 180.** According to Maxwell's law of distribution of velocities of molecules, the most probable velocity is :-
- greater than the mean speed
 - equal to the mean speed
 - equal to the root mean square speed
 - less than the root mean square speed

181. The ratio of average translational kinetic energy to rotational kinetic energy of a diatomic molecule at temperature T is

- (1) 3 (2) $7/5$
(3) $5/3$ (4) $3/2$

182. For hydrogen gas $c_p - c_v = a$ and for oxygen gas $c_p - c_v = b$ then the relation between a and b is (where c_p & c_v are gram specific heats)

- (1) $a = 16b$ (2) $b = 16a$
(3) $a = b$ (4) None of these

183. A gas mixture consists of 2 moles of oxygen and 4 moles of argon at temperature T . Neglecting all vibrational modes, the total internal energy of the system is

- (1) $4RT$ (2) $15RT$
(3) $9RT$ (4) $11RT$

184. The average kinetic energy of a gas molecule at 27°C is 6.21×10^{-21} J. Its average kinetic energy at 227°C will be

- (1) 52.2×10^{-21} J
(2) 5.22×10^{-21} J
(3) 10.35×10^{-21} J
(4) 11.35×10^{-21} J

185. Two containers A and B contain molecular gas at same temperature with masses of molecules are m_A and m_B , then relation of momentum P_A and P_B will be-

- (1) $P_A = P_B$
(2) $P_A = \left(\frac{m_A}{m_B}\right)^{1/2} P_B$
(3) $P_A = \left(\frac{m_B}{m_A}\right)^{1/2} P_B$
(4) $P_A = \left(\frac{m_A}{m_B}\right) P_B$

186. A cylinder of 200 litre capacity is containing H_2 . The total translational kinetic energy of molecules is 1.52×10^5 J. The pressure of H_2 in the cylinder will be in N m^{-2} :-

- (1) 2×10^5 (2) 3×10^5
(3) 4×10^5 (4) 5×10^5

(187-188) Five moles of helium are mixed with two moles of hydrogen to form a mixture. Take molar mass of helium $M_1 = 4\text{g}$ and that of hydrogen $M_2 = 2\text{g}$

187. The equivalent molar mass of the mixture is

- (1) 6g (2) $\frac{13\text{g}}{7}$ (3) $\frac{18\text{g}}{7}$ (4) $\frac{24\text{g}}{7}$

188. The equivalent value of γ in the above question is

- (1) 1.59 (2) 1.53
(3) 1.56 (4) none

189. Two monoatomic ideal gas at temperature T_1 and T_2 are mixed. There is no loss of energy. If the mass of molecules of the two gases are m_1 and m_2 and number of their molecules are n_1 and n_2 respectively, then temperature of the mixture will be :

- (1) $\frac{T_1 + T_2}{n_1 + n_2}$ (2) $\frac{T_1}{n_1} + \frac{T_2}{n_2}$
(3) $\frac{n_2 T_1 + n_1 T_2}{n_1 + n_2}$ (4) $\frac{n_1 T_1 + n_2 T_2}{n_1 + n_2}$

190. The total kinetic energy of 1 mole of N_2 at 27°C will be approximately :-

- (1) 1500 J
(2) 1500 calorie
(3) 1500 kilo calorie
(4) 1500 erg.

191. Mean kinetic energy (or average energy) per gm. molecule of a monoatomic gas is given by :

- (1) $3RT/2$ (2) $kT/2$
(3) $RT/3$ (4) $3kT/2$

192. Relation between the ratio of specific heats (γ) of gas and degree of freedom 'f' will be

- (1) $\gamma = f + 2$ (2) $\frac{1}{\gamma} = \frac{1}{f} + \frac{1}{2}$
(3) $f = 2 / (\gamma - 1)$ (4) $f = 2(\gamma - 1)$

193. Relation between pressure (P) and energy density (E) of an ideal gas is -

- (1) $P = \frac{2E}{3}$ (2) $P = \frac{3E}{2}$
(3) $P = \frac{3E}{5}$ (4) $P = E$

194. On mixing 1 g mole of a monoatomic with 1 g mole of a diatomic gas the specific heat of mixture at constant volume will be :-
 (1) R (2) $3/2 R$ (3) $2R$ (4) $5/2R$

195. Absolute zero temperature is one at which-
 (1) All liquids convert into solid
 (2) All gases convert to solid
 (3) All matter is in solid state
 (4) The K.E. of molecules becomes zero

196. For a gas $\frac{R}{C_v} = 0.67$. This gas is made up of molecules which are :
 (1) Diatomic
 (2) Mixture of diatomic and polyatomic molecules
 (3) Monoatomic
 (4) Polyatomic

197. If the total number of H_2 molecules is double that of the O_2 molecules then ratio of total kinetic energies of H_2 to that of O_2 at 300 K is :
 (1) 1 : 1 (2) 1 : 2
 (3) 2 : 1 (4) 1 : 16

198. At which of the following temperature any gas has average molecular kinetic energy double that of at $20^\circ C$
 (1) $40^\circ C$ (2) $80^\circ C$
 (3) $313^\circ C$ (4) $586^\circ C$

199. When temperature is increased from $0^\circ C$ to $273^\circ C$, in what ratio of final to initial the average kinetic energy of molecules change ?
 (1) 1 (2) 3 (3) 4 (4) 2

200. The kinetic energy associated with per degree of freedom of a molecule is -
 (1) $\frac{1}{2} MC_{rms}^2$ (2) kT
 (3) $kT / 2$ (4) $3 kT / 2$

201. If 2 gm moles of a diatomic gas and 1 gm mole of a mono-atomic gas are mixed then the value of γ (C_p / C_v) for mixture will be :-
 (1) $\frac{13}{19}$ (2) $\frac{19}{13}$ (3) $\frac{7}{5}$ (4) $\frac{5}{3}$

202. Which of the following statement is true according to kinetic theory of gases ?

- (1) The collision between two molecules is inelastic and the time between two collisions is less than the time taken during the collision.
- (2) There is a force of attraction between the molecules
- (3) All the molecules of a gas move with same velocity
- (4) The average of the distances travelled between two successive collisions is mean free path.

203. Gas exerts pressure on the walls of container because the molecules-
 (1) Are losing their Kinetic energy
 (2) Are getting stuck to the walls
 (3) Are transferring their momentum to walls
 (4) Are accelerated towards walls.

204. For a diatomic gas, change in internal energy for unit change in temperature at constant pressure and volume is U_1 and U_2 respectively then $U_1 : U_2$ is :
 (1) 5 : 3 (2) 7 : 5
 (3) 1 : 1 (4) 5 : 7

205. The specific heat of an ideal gas depends on temperature is -
 (1) $\frac{1}{T}$
 (2) T
 (3) \sqrt{T}
 (4) Does not depend on temperature

206. The specific heat of a gas :
 (1) Has only two values C_p and C_v
 (2) Has a unique value at a given temperature
 (3) Can have any value between 0 and ∞
 (4) Depends upon the mass of the gas

207. 22 g of CO_2 at $27^\circ C$ is mixed with 16 g of O_2 at $37^\circ C$. The temperature of the mixture is :-
 (At room temperature, degrees of freedom of $CO_2 = 7$ and degrees of freedom of $O_2 = 5$)
 (1) $31.16^\circ C$ (2) $27^\circ C$
 (3) $37^\circ C$ (4) $30^\circ C$

208. At 27°C temperature, the kinetic energy of an ideal gas is E_1 . If the temperature is increased to 327°C , then kinetic energy would be

- (1) $2 E_1$ (2) $\frac{1}{2} E_1$
 (3) $\sqrt{2} E_1$ (4) $\frac{1}{\sqrt{2}} E_1$

209. Oxygen and hydrogen gases are at temperature T . Then K.E of molecules of oxygen gas is equal to how many times of average K.E. of molecules of hydrogen gas :-

- (1) 16 times (2) 8 times
 (3) Equal (4) $1/16$ times

210. The average energy of the molecules of a monoatomic gas at temperature T is :-
 (K = Boltzmann constant)

- (1) $\frac{1}{2} kT$ (2) kT
 (3) $\frac{3}{2} kT$ (4) $\frac{5}{2} kT$

211. A diatomic molecule has

- (1) 1 degree of freedom
 (2) 3 degree of freedom
 (3) 5 degree of freedom
 (4) 6 degree of freedom

212. Two moles of monoatomic gas are mixed with 1 mole of a diatomic gas. Then γ for the mixture is:

- (1) 1.4 (2) 1.55 (3) 1.62 (4) 1.67

ZEROth AND FIRST LAW OF THERMODYNAMICS, HEAT, WORK AND INTERNAL ENERGY

213. The first law of thermodynamics is based on :-

- (1) Law of conservation of energy
 (2) Law of conservation of mechanical energy
 (3) Law of conservation of gravitational P.E.
 (4) None of the above

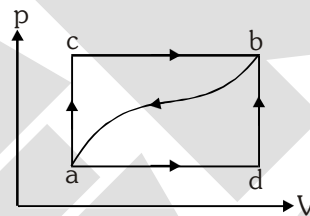
214. In a process, 500 calories of heat is given to a system and at the same time 100 joules of work is done on the system. The increase in the internal energy of the system is :-

- (1) 40 calories (2) 1993 joules
 (3) 2193 joules (4) 82 calories

215. In a thermodynamic process pressure of a fixed mass of a gas is changed in such a manner that the gas releases 20 joules of heat and 8 joules of work was done on the gas. If the initial internal energy of the gas was 30 joules, then the final internal energy will be:-

- (1) 2 J (2) 42 J (3) 18 J (4) 58 J

216. When a system is taken from state 'a' to state 'b' along the path 'acb', it is found that a quantity of heat $Q = 200 \text{ J}$ is absorbed by the system and a work $W = 80 \text{ J}$ is done by it. Along the path 'adb', $Q = 144 \text{ J}$. The work done along the path 'adb' is



- (1) 6 J (2) 12 J (3) 18 J (4) 24 J

217. In the above question, if the work done on the system along the curved path 'ba' is 52 J, heat absorbed is

- (1) -140 J (2) -172 J
 (3) 140 J (4) 172 J

218. In question No. 217, if $U_a = 40 \text{ J}$, value of U_b will be

- (1) -50 J (2) 100 J
 (3) -120 J (4) 160 J

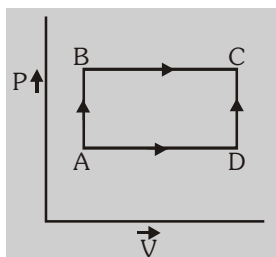
219. In question No. 217, if $U_d = 88 \text{ J}$, heat absorbed for the path 'db' is

- (1) -72 J (2) 72 J
 (3) 144 J (4) -144 J

220. 1 kg of a gas does 20 kJ of work and receives 16 kJ of heat when it is expanded between two states. A second kind of expansion can be found between the same initial and final state which requires a heat input of 9 kJ. The work done by the gas in the second expansion is :

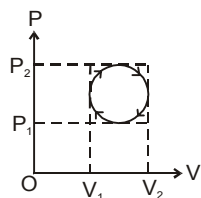
- (1) 32 kJ (2) 5 kJ
 (3) -4 kJ (4) 13 kJ

- 221.** As shown in the figure the amount of heat absorbed along the path ABC is 90J and the amount of work done by the system is 30 J. If the amount of work done along the path ADC is 20 J then amount of heat absorbed will be :-



- (1) 80 J (2) 90 J (3) 110 J (4) 120 J

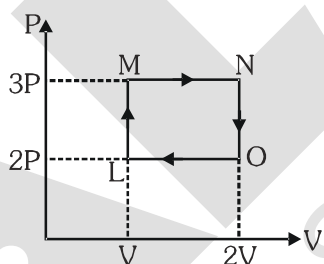
- 222.** In a cyclic process shown on the P – V diagram, the magnitude of the work done is :



- (1) $\pi \left(\frac{P_2 - P_1}{2} \right)^2$ (2) $\pi \left(\frac{V_2 - V_1}{2} \right)^2$
 (3) $\frac{\pi}{4} (P_2 - P_1) (V_2 - V_1)$ (4) $\pi (P_2 V_2 - P_1 V_1)$

- 223.** The work by an ideal monoatomic gas along the cyclic path LMNOL is

- (1) PV
 (2) 2PV
 (3) 3 PV
 (4) 4 PV



- 224.** For a gas $C_v = 4.96$ cal/mole K, the increase in internal energy of 2 mole gas in heating from 340 K to 342 K will be :-

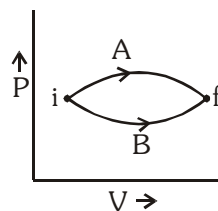
- (1) 27.80 cal (2) 19.84 cal
 (3) 13.90 cal (4) 9.92 cal

- 225.** When a system changes from one to another state the value of work done :-

- (1) Depends on the force acting on the system
 (2) Depends on the nature of material present in a system
 (3) Does not depend on the path
 (4) Depends on the path

- 226.** A system is taken along the paths A and B as shown. If the amounts of heat given in these processes are Q_A and Q_B and change in internal energy are ΔU_A and ΔU_B respectively then :-

- (1) $Q_A = Q_B$; $\Delta U_A < \Delta U_B$
 (2) $Q_A \geq Q_B$; $\Delta U_A = \Delta U_B$
 (3) $Q_A < Q_B$; $\Delta U_A > \Delta U_B$
 (4) $Q_A > Q_B$; $\Delta U_A = \Delta U_B$



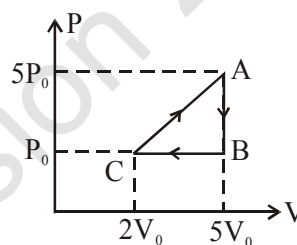
- 227.** If the heat of 110 J is added to a gaseous system and change in internal energy is 40 J, then the amount of external work done is :

- (1) 180 J (2) 70 J (3) 110 J (4) 30 J

- 228.** If amount of heat supplied is Q , work done is W and change in internal energy is $mC_v dT$, then relation among them is. (C_v = gram specific heat)

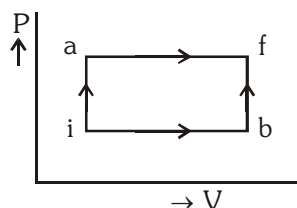
- (1) $mC_v dT = Q + W$
 (2) $Q = W + mC_v dT$
 (3) $Q + mC_v dT = W$
 (4) None of these.

- 229.** The work done by a gas taken through the closed process ABCA is



- (1) $6P_0V_0$ (2) $4P_0V_0$ (3) P_0V_0 (3) zero

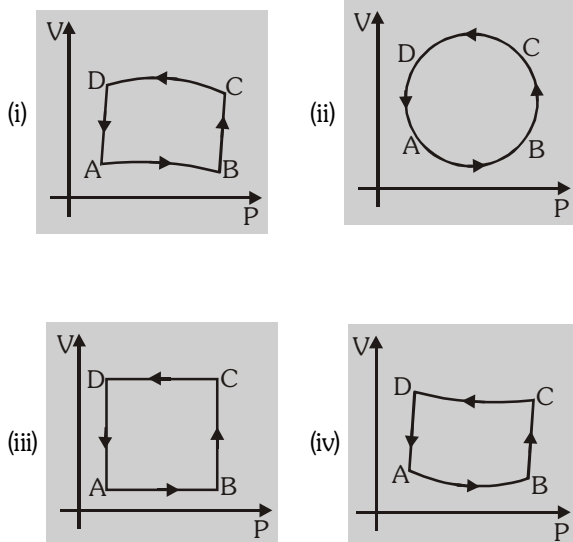
- 230.** In the given figure, the initial and final states of a gas are shown by points i and f. The internal energy of the gas at i is 10J. For the path iaf : $dQ = 50$ J and $dW = 20$ J. For the path ibf, if $dQ = 36$ J the value of dW will be equal to



- (1) 4J (2) 12J (3) 18J (4) 6J

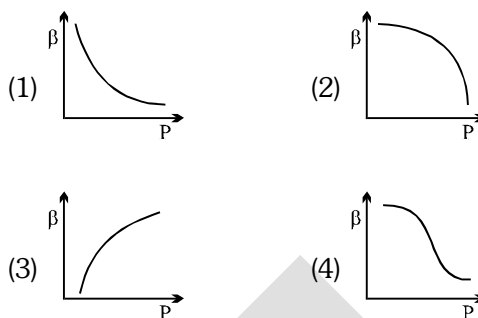
CYCLIC, ISOCHORIC, ISOBARIC, ISOTHERMAL, ADIABATIC AND POLYTROPIC PROCESS

- 231.** In the diagrams (i) to (iv) of variation of volume with changing pressure is shown. A gas is taken along the path ABCDA. The change in internal energy of the gas will be:-



- (1) Positive in all cases (i) to (iv)
 (2) Positive in cases (i), (ii) and (iii) but zero in case (iv)
 (3) Negative in cases (i), (ii) and (iii) but zero in case (iv)
 (4) Zero in all the four cases
- 232.** The temperature of 5 moles of a gas which was held at constant volume was changed from 100°C to 120°C . The change in internal energy was found to be 80 joules. The total heat capacity of the gas at constant volume will be equal to:-
 (1) 8 J/K (2) 0.8 J/K
 (3) 4.0 J/K (4) 0.4 J/K
- 233.** Monoatomic, diatomic and triatomic gases whose initial volume and pressure are same, are compressed till their volume become half the initial volume.
 (1) If the compression is adiabatic then monoatomic gas will have maximum final pressure.
 (2) If the compression is adiabatic then triatomic gas will have maximum final pressure.
 (3) If the compression is adiabatic then their final pressure will be same.
 (4) If the compression is isothermal then their final pressure will be different.

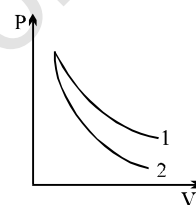
- 234.** Which of the following graphs correctly represents the variation of $\beta = -(dV/dP)/V$ with P for an ideal gas at constant temperature?



- 235.** The adiabatic Bulk modulus of a diatomic gas at atmospheric pressure is
 (1) 0 Nm^{-2}
 (2) 1 Nm^{-2}
 (3) $1.4 \times 10^4 \text{ Nm}^{-2}$
 (4) $1.4 \times 10^5 \text{ Nm}^{-2}$

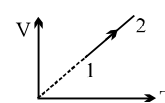
- 236.** A given quantity of an ideal gas is at pressure P and absolute temperature T. The isothermal bulk modulus of the gas is :
 (1) $2P/3$ (2) P (3) $3P/2$ (4) $2P$

- 237.** P-V plots for two gases during adiabatic processes are shown in the figure. Plots 1 and 2 should correspond respectively to



- (1) He and O_2 (2) O_2 and He
 (3) He and Ar (4) O_2 and N_2

- 238.** An ideal gas undergoes the process $1 \rightarrow 2$ as shown in the figure, the heat supplied and work done in the process is ΔQ and ΔW respectively. The ratio $\Delta Q : \Delta W$ is



- (1) $\gamma : \gamma - 1$ (2) γ
 (3) $\gamma - 1$ (4) $\gamma - 1 : \gamma$

239. For an adiabatic expansion of a perfect gas, the value of $\Delta P/P$ is equal to:-

- (1) $-\sqrt{\gamma} \Delta V/V$ (2) $-\Delta V/V$
(3) $-\gamma \Delta V/V$ (4) $-\gamma^2 \Delta V/V$

240. An ideal gas at 27°C is compressed adiabatically to $8/27$ of its original volume. If $\gamma = 5/3$, then the rise in temperature is:-

- (1) 450 K (2) 375 K (3) 675 K (4) 405 K

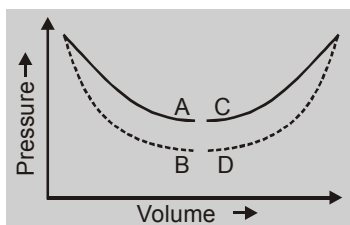
241. For monoatomic gas the relation between pressure of a gas and temperature T is given by $P \propto T^C$. Then value of C will be : (For adiabatic process)

- (1) $\frac{5}{3}$ (2) $\frac{2}{5}$ (3) $\frac{3}{5}$ (4) $\frac{5}{2}$

242. A gas for which $\gamma = 5/3$ is heated at constant pressure. The percentage of total heat given that will be used for external work is :

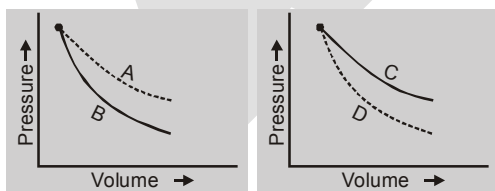
- (1) 40% (2) 30% (3) 60% (4) 20%

243. In which of the figure no heat exchange between the gas and the surroundings will take place, if the gas is taken along curve:
(curves are isothermal and adiabatic)



- (1) A (2) B (3) C (4) D

244. In the following figures, four curves A, B, C, D are shown the curves are :-



- (1) Isothermal for A and B while adiabatic for C and D
(2) Isothermal for A and C while adiabatic for B and D
(3) Isothermal for A and D
(4) Adiabatic for A and C while isothermal for B and D

245. Equal volumes of a perfect gas are compressed to half of their initial volumes. The first is brought about by isothermal process and the second by adiabatic process then :

- (1) Both temperature and pressure will increase in the isothermal process.
(2) In the isothermal process, the temperature will decrease and pressure will increase
(3) Both temperature and pressure will increase in adiabatic process
(4) In the adiabatic process, the temperature will decrease and pressure will increase

246. A vessel contains an ideal monoatomic gas which expands at constant pressure, when heat Q is given to it. Then the work done in expansion is :

- (1) Q (2) $\frac{3}{5}Q$ (3) $\frac{2}{5}Q$ (4) $\frac{2}{3}Q$

247. One mole of an ideal gas at temperature T_1 expands according to the law $\frac{P}{V^2} = a$ (constant). The work done by the gas till temperature of gas becomes T_2 is :

- (1) $\frac{1}{2}R(T_2 - T_1)$ (2) $\frac{1}{3}R(T_2 - T_1)$
(3) $\frac{1}{4}R(T_2 - T_1)$ (4) $\frac{1}{5}R(T_2 - T_1)$

248. When an ideal diatomic gas is heated at constant pressure, the fraction of the heat energy supplied which increases the internal energy of the gas is -

- (1) $2/5$ (2) $3/5$ (3) $3/7$ (4) $5/7$

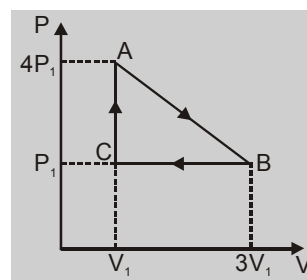
249. Pressure-temperature relationship for an ideal gas undergoing adiabatic change is ($\gamma = C_p/C_v$)

- (1) $PT^\gamma = \text{constant}$
(2) $PT^{-1+\gamma} = \text{constant}$
(3) $PT^{\gamma-1}T^\gamma = \text{constant}$
(4) $P^{1-\gamma}T^\gamma = \text{constant}$

250. The value of internal energy in an adiabatic process :-

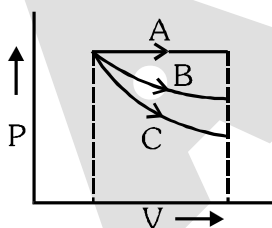
- (1) Remains unchanged
(2) Only increases
(3) Only diminishes
(4) May diminish and may also increase

- 251.** One mole of an ideal monoatomic gas is heated at a constant pressure of one atmosphere from 0°C to 100°C . Then the change in the internal energy is
 (1) $20.80 \times 10^2 \text{ J}$
 (2) $12.48 \times 10^2 \text{ J}$
 (3) $832 \times 10^2 \text{ J}$
 (4) $6.25 \times 10^2 \text{ J}$
- 252.** The specific heat of a gas at constant pressure is more than that of the same gas at constant volume because :-
 (1) Work is done in the expansion of gas at constant pressure
 (2) Work is done in the expansion of the gas at constant volume
 (3) The molecular attraction increase under constant pressure
 (4) The vibration of molecules increases under constant pressure
- 253.** When a gas is adiabatically compressed then its temperature increase because :-
 (1) Work done is zero
 (2) Internal energy is increased
 (3) Heat is supplied to it
 (4) No change in pressure
- 254.** Air is filled in a tube of the wheel of a car at 27°C and 2 atm pressure if the tube is suddenly bursts, the final temperature of air will be :-
 ($\gamma = 1.5$, $2^{1/3} = 1.251$)
 (1) -33°C (2) 0°C
 (3) 21.6°C (4) 240°C
- 255.** Specific heat of a gas undergoing adiabatic change is :
 (1) Zero (2) Infinite
 (3) Positive (4) Negative
- 256.** A quantity of air ($\gamma = 1.4$) at 27°C is compressed suddenly, the temperature of the air system will :
 (1) Fall
 (2) Rise
 (3) Remain unchanged
 (4) First rise and then fall
- 257.** A gas specimen in one vessel is expanded isothermally to double its volume and a similar specimen in the second vessel is expanded adiabatically the same extent, then :
 (1) In the second vessel, both pressure and work done are more
 (2) In the second vessel, pressure is more, but the work done is less.
 (3) In the first vessel, both pressure & work done are more.
 (4) In the first vessel, pressure is more, but work done is less
- 258.** If an ideal gas is compressed during isothermal process then :-
 (1) No work is done against gas
 (2) heat is rejected by gas
 (3) Its internal energy will increase
 (4) Pressure does not change
- 259.** Graphs between P-V diagram for isothermal and adiabatic processes are drawn the relation between their slopes will be :-
 (1) Slope of adiabatic curve $= \gamma$ (slope of isothermal curve)
 (2) Slope of isothermal curve $= \gamma$ (slope of adiabatic curve)
 (3) Slope of isothermal curve = slope of adiabatic curve
 (4) Slope of adiabatic curve $= \gamma^2$ (slope of isothermal curve)
- 260.** An ideal gas is taken round the cycle ABCA. In the cycle the amount of work done involved is :-



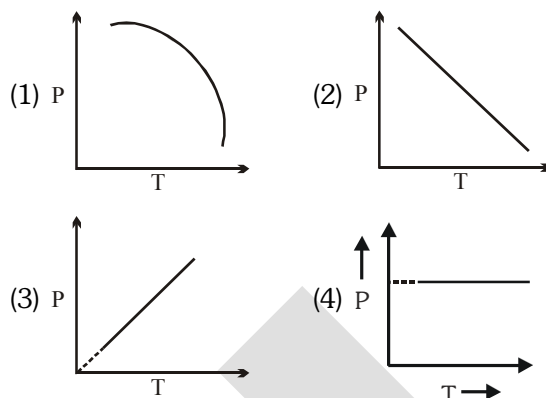
- (1) $12 P_1 V_1$ (2) $6 P_1 V_1$
 (3) $3 P_1 V_1$ (4) $P_1 V_1$

- 261.** One mole ideal gas is compressed adiabatically at 27°C . Its temperature becomes 102°C . The work done in this process will be :- ($\gamma = 1.5$)
- (1) -625 J (2) 625 J
(3) 1245 J (4) -1245 J
- 262.** In an isometric change :
- (1) $\delta Q = dU$ (2) $\delta W = dU$
(3) $\delta Q + \delta W = dU$ (4) None of these
- 263.** The volume of a gas expands by 0.25 m^3 at a constant pressure of 10^3 N/m^2 . The work done is equal to
- (1) 2.5 erg (2) 250 J
(3) 250 W (4) 250 N
- 264.** The process in which the heat given to a system is completely transformed into work is for ideal gas :-
- (1) Isobaric process (2) Isometric process
(3) Isothermal process (4) Adiabatic process
- 265.** The volume of a poly-atomic gas $\left(\gamma = \frac{4}{3}\right)$ compressed adiabatically to $\frac{1}{8^{\text{th}}}$ of the original volume. If the original pressure of the gas is P_0 the new pressure will be :
- (1) $8 P_0$ (2) $16 P_0$
(3) $6 P_0$ (4) $2 P_0$
- 266.** During isothermal, isobaric and adiabatic processes, work done for same change in volume will be maximum for :-

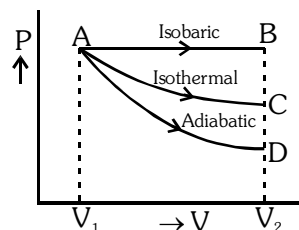


- (1) Isothermal (2) Isobaric
(3) Adiabatic (4) None of the above
- 267.** In an adiabatic process the quantity which remains constant is :-
- (1) Temperature
(2) Pressure
(3) Total heat content of the system
(4) Volume

- 268.** Graph of isometric process is :-

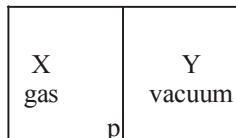


- 269.** An ideal gas undergoes an adiabatic change in volume (V) with pressure (P). Then :-
- (1) $P^{\gamma}V = \text{constant}$ (2) $PV^{\gamma} = \text{constant}$
(3) $(PV)^{\gamma} = \text{constant}$ (4) $PV = \text{constant}$
- 270.** 300 calories of heat is supplied to raise the temperature of 50 gm of air from 20°C to 30°C without any change in its volume. Change in internal energy per gram of air is
- (1) zero (2) 0.6 calories
(3) 1.2 calories (4) 6.0 calories
- 271.** A gas is expanded from volume V_1 to volume V_2 in three processes, shown in the figure. If U_A , U_B , U_C and U_D represent the internal energies of the gas in state A, B, C and D respectively, the which of the following is not correct



- $$\begin{array}{ll} (1) U_B - U_A > 0 & (2) U_C - U_A = 0 \\ (3) U_D - U_A < 0 & (4) U_B = U_C = U_D \end{array}$$

- 272.** A closed container is fully insulated from outside. One half of it is filled with an ideal gas X separated by a plate P from the other half Y which contains a vacuum as shown in figure. When P is removed, X moves into Y. Which of the following statements is correct?

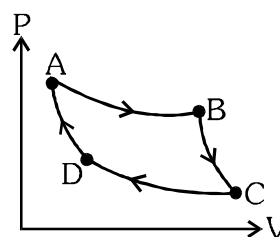


- (1) No work is done by X
- (2) X decreases in temperature
- (3) X increases in internal energy
- (4) X doubles in pressure

SECOND LAW OF THE THERMODYNAMICS, HEAT ENGINES AND REFRIGERATORS

- 273.** According to the second law of thermodynamics :
- (1) heat energy cannot be completely converted to work
 - (2) work cannot be completely converted to heat energy
 - (3) for all cyclic processes we have $dQ/T < 0$
 - (4) the reason all heat engine efficiencies are less than 100% is friction, which is unavoidable
- 274.** "Heat cannot flow by itself from a body at lower temperature to a body at higher temperature" is a statement or consequence of :
- (1) second law of thermodynamics
 - (2) conservation of momentum
 - (3) conservation of mass
 - (4) first law of thermodynamics
- 275.** A Carnot engine takes 3×10^6 cal of heat from reservoir at 627°C and gives it to a sink at 27°C . Then work done by the engine is
- (1) 4.2×10^6 J
 - (2) 8.4×10^6 J
 - (3) 16.8×10^6 J
 - (4) zero
- 276.** A reversible refrigerator operates between a low temperature reservoir at T_C and a high temperature reservoir at T_H . Its coefficient of performance is given by :
- (1) $(T_H - T_C)/T_C$
 - (2) $T_C/(T_H - T_C)$
 - (3) $(T_H - T_C)/T_H$
 - (4) $T_H/(T_H - T_C)$

- 277.** In the given graph the isothermal curves are :-



- (1) AB and CD
- (2) AB and AD
- (3) AD and BC
- (4) BC and CD

- 278.** In the above question, the curve for which the heat is absorbed from the surroundings is :-
- (1) AB
 - (2) BC
 - (3) CD
 - (4) DA

- 279.** A carnot engine shows efficiency of 40% on taking energy at 500 K. To increase the efficiency to 50%, at what temperature it should take energy ?
- (1) 400 K
 - (2) 700 K
 - (3) 600 K
 - (4) 800 K

- 280.** A Carnot engine, whose efficiency is 40%, takes in heat from a source maintained at a temperature of 500 K. It is desired to have an engine of efficiency 60%. Then, the intake temperature for the same exhaust (sink) temperature must be :
- (1) 750 K
 - (2) 600 K
 - (3) Efficiency of Carnot engine cannot be made larger than 50%
 - (4) 1200 K

- 281.** If the system takes 100 cal. heat, and releases 80 cal to sink, if source temperature is 127°C find the sink temperature :-
- (1) 47°C
 - (2) 127°C
 - (3) 67°C
 - (4) 107°C

- 282.** A carnot engine working between 300 K and 600 K has work out put of 800 J per cycle. The amount of heat energy supplied to the engine from source per cycle will be :
- (1) 800 J
 - (2) 1600 J
 - (3) 1200 J
 - (4) 900 J

- 283.** The efficiency of carnot engine is 50% and temperature of sink is 500K. If temperature of source is kept constant and its efficiency raised to 60%, then the required temperature of the sink will be :-
- (1) 100 K
 - (2) 600 K
 - (3) 400 K
 - (4) 500 K

284. An ideal gas heat engine operates in carnot cycle between 227°C and 127°C . It absorbs 6×10^4 cal of heat at higher temperature. Then amount of heat converted to work is :

- (1) 2.4×10^4 cal (2) 6×10^4 cal
(3) 1.2×10^4 cal (4) 4.8×10^4 cal

285. A refrigerator works between temperature -10°C and 27°C , the coefficient of performance is :

- (1) 7.1 (2) 1
(3) 8.1 (4) 15.47

EXERCISE-I (Conceptual Questions)

ANSWER KEY

Que.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15
Ans.	1	1	1	2	4	3	3	1	2	2	2	1	3	4	1
Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	3	1	1	1	1	4	2	4	1	1	3	3	4	3	1
Que.	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45
Ans.	4	4	3	1	2	3	2	2	2	3	1	1	1	2	1
Que.	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60
Ans.	4	4	2	1	3	2	2	2	4	2	3	1	1	2	2
Que.	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75
Ans.	4	3	3	1	3	3	3	1	3	1	4	3	4	1	2
Que.	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90
Ans.	2	3	2	4	3	3	3	3	4	2	4	2	4	4	2
Que.	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105
Ans.	1	1	1	2	3	3	1	3	2	2	3	1	4	1	3
Que.	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120
Ans.	4	4	3	2	4	3	1	1	2	3	3	2	1	2	2
Que.	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135
Ans.	4	1	4	4	4	4	4	3	1	3	3	3	4	3	2
Que.	136	137	138	139	140	141	142	143	144	145	146	147	148	149	150
Ans.	3	3	2	3	2	4	4	3	4	3	4	3	4	4	2
Que.	151	152	153	154	155	156	157	158	159	160	161	162	163	164	165
Ans.	1	3	3	3	2	1	2	1	1	3	2	2	2	2	1
Que.	166	167	168	169	170	171	172	173	174	175	176	177	178	179	180
Ans.	1	1	4	3	1	4	1	4	2	4	3	4	1	2	4
Que.	181	182	183	184	185	186	187	188	189	190	191	192	193	194	195
Ans.	4	1	4	3	2	4	4	3	4	2	1	3	1	3	4
Que.	196	197	198	199	200	201	202	203	204	205	206	207	208	209	210
Ans.	3	3	3	4	3	2	4	3	3	4	3	1	1	3	3
Que.	211	212	213	214	215	216	217	218	219	220	221	222	223	224	225
Ans.	3	2	1	3	3	4	2	4	2	4	1	3	1	2	4
Que.	226	227	228	229	230	231	232	233	234	235	236	237	238	239	240
Ans.	4	2	2	1	4	4	3	1	1	4	2	2	1	3	2
Que.	241	242	243	244	245	246	247	248	249	250	251	252	253	254	255
Ans.	4	1	2	2	3	3	2	4	4	4	2	1	2	1	1
Que.	256	257	258	259	260	261	262	263	264	265	266	267	268	269	270
Ans.	2	3	2	1	3	4	1	2	3	2	2	3	3	2	4
Que.	271	272	273	274	275	276	277	278	279	280	281	282	283	284	285
Ans.	4	1	1	1	2	2	1	1	3	1	1	2	3	3	1

EXERCISE-II (Previous Year Questions)

AIPMT/NEET & AIIMS (2006-2018)

AIPMT 2006

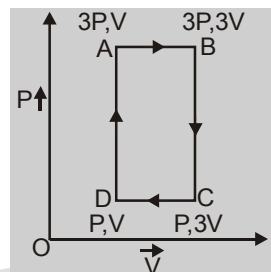
- A black body emits radiation of maximum intensity at 5000 \AA when its temperature is 1227°C . If its temperature is increased by 1000°C then the maximum intensity of emitted radiation will be at:
 - 2754.8 \AA
 - 3000 \AA
 - 3500 \AA
 - 4000 \AA
- The translational kinetic energy of molecules of one mole of a monoatomic gas is $U = 3NkT/2$. The value of molar specific heat of gas under constant pressure will be :
 - $\frac{3}{2}R$
 - $\frac{5}{2}R$
 - $\frac{7}{2}R$
 - $\frac{9}{2}R$
- The molar specific heat at constant pressure of an ideal gas is $(7/2)R$. The ratio of specific heat at constant pressure to that at constant volume is :
 - $\frac{7}{5}$
 - $\frac{8}{7}$
 - $\frac{5}{7}$
 - $\frac{9}{7}$

AIIMS 2006

- Three objects coloured black, gray and white can withstand hostile conditions upto 2800°C . These objects are thrown into a furnace where each of them attains a temperature of 2000°C . Which object will glow brightest :-
 - The white object
 - The black object
 - All glow with equal brightness
 - Gray object
- Two balloons are filled, one with pure He gas and the other by air, respectively. If the pressure and temperature of these balloons are same then the number of molecules per unit volume is :-
 - more in the He filled balloon
 - same in both balloons
 - more in air filled balloon
 - in the ratio of 1 : 4

AIPMT 2007

- An ideal monoatomic gas is taken round the cycle ABCDA as shown in following P - V diagram. The work done during the cycle is :



- PV
 - 2 PV
 - 4 PV
 - Zero
- The $\left(\frac{W}{Q}\right)$ of a carnot-engine is $\frac{1}{6}$, now the temperature of sink is reduced by 62°C , then this ratio becomes twice, therefore the initial temperature of the sink and source are respectively:-
 - $33^\circ \text{C}, 67^\circ \text{C}$
 - $37^\circ \text{C}, 99^\circ \text{C}$
 - $67^\circ \text{C}, 33^\circ \text{C}$
 - $97^\circ \text{C}, 37^\circ \text{C}$

AIPMT 2008

- A new scale of temperature (which is linear) called the W scale, the freezing and boiling points of water are 39°W and 239°W respectively. What will be the temperature on the new scale, corresponding to a temperature of 39°C on the Celsius scale ?
 - 200°W
 - 139°W
 - 78°W
 - 117°W
- At 10°C the value of the density of a fixed mass of an ideal gas divided by its pressure is x. At 110°C this ratio is :-
 - $\frac{10}{110}x$
 - $\frac{283}{383}x$
 - x
 - $\frac{383}{283}x$
- If Q, E and W denote the heat added, change in internal energy and the work done respectively in a closed cycle process, then :-
 - $E = 0$
 - $Q = 0$
 - $W = 0$
 - $Q = W = 0$

AIPMT 2009

11. The two ends of a rod of length L and a uniform cross sectional area A are kept at two temperatures T_1 and T_2 ($T_1 > T_2$). The rate of heat transfer $\frac{dQ}{dt}$, through the rod in a steady state is given by :-

(1) $\frac{dQ}{dt} = \frac{kA(T_1 - T_2)}{L}$ (2) $\frac{dQ}{dt} = \frac{kL(T_1 - T_2)}{A}$

(3) $\frac{dQ}{dt} = \frac{k(T_1 - T_2)}{LA}$ (4) $\frac{dQ}{dt} = kLA(T_1 - T_2)$

12. A black body, at a temperature of 227°C radiates heat at a rate of $7 \text{ cal cm}^{-2} \text{ s}^{-1}$. At a temperature of 727°C , the rate of heat radiated in the same units will be :-

- (1) 80 (2) 60 (3) 50 (4) 112

13. In thermodynamic processes which of the following statement is not true :-

- (1) In an adiabatic process $PV^\gamma = \text{constant}$
 (2) In an adiabatic process the system is insulated from the surroundings
 (3) In an isochoric process pressure remains constant
 (4) In an Isothermal process the temperature remains constant

14. The change in internal energy in a system that has absorbed 2 Kcals of heat and 500 J of work done is :-

- (1) 7900J (2) 8900J
 (3) 6400J (4) 5400J

AIPMT (Pre) 2010

15. A cylindrical metallic rod in thermal contact with two reservoirs of heat at its two ends conducts an amount of heat Q in time t . The metallic rod is melted and the material is formed into a rod of half the radius of the original rod. What is the amount of heat conducted by the new rod, when placed in thermal contact with the two reservoirs in time t ?

- (1) $\frac{Q}{2}$ (2) $\frac{Q}{4}$
 (3) $\frac{Q}{16}$ (4) $2Q$

16. Total radiant energy per unit area, per unit time normal to the direction of incidence, received at a distance R from the centre of a star of radius r , whose outer surface radiates as a black body at a temperature T Kelvin is given by :-

- (1) $\frac{4\pi\sigma r^2 T^4}{R^2}$ (2) $\frac{\sigma r^2 T^4}{R^2}$
 (3) $\frac{\sigma r^2 T^4}{4\pi r^2}$ (4) $\frac{\sigma r^4 T^4}{r^4}$

(Where σ is Stefan's Constant)

17. If ΔU and ΔW represent the increase in internal energy and work done by the system respectively in a thermodynamic process, which of the following is true ?

- (1) $\Delta U = -\Delta W$, in an isothermal process
 (2) $\Delta U = -\Delta W$, in an adiabatic process
 (3) $\Delta U = \Delta W$, in an isothermal process
 (4) $\Delta U = \Delta W$, in an adiabatic process

AIPMT (Mains) 2010

18. If c_p and c_v denote the specific heats (per unit mass) of an ideal gas of molecular weight M , then :-

- (1) $c_p - c_v = R$ (2) $c_p - c_v = \frac{R}{M}$
 (3) $c_p - c_v = MR$ (4) $c_p - c_v = \frac{R}{M^2}$

where R is the molar gas constant

19. A monoatomic gas at pressure P_1 and volume V_1 is compressed adiabatically to $1/8^{\text{th}}$ its original volume. What is the final pressure of the gas :-

- (1) P_1 (2) $16P_1$ (3) $32P_1$ (4) $64P_1$

AIPMT (Pre) 2011

20. During an isothermal expansion, a confined ideal gas does +150 J of work against its surroundings. This implies that :-

- (1) 150 J of heat has been removed from the gas
 (2) 300 J of heat has been added to the gas
 (3) No heat is transferred because the process is isothermal
 (4) 150 J of heat has been added to the gas

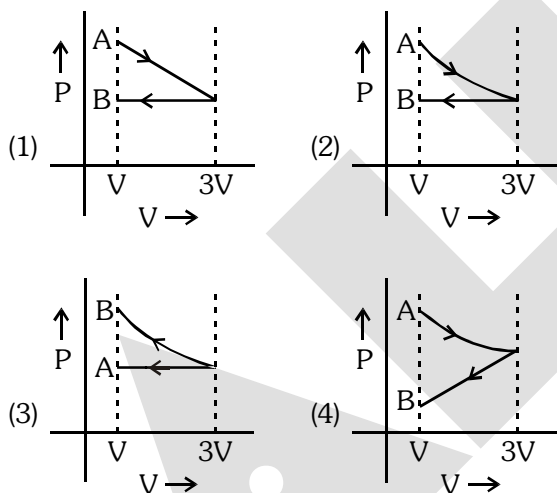
21. A mass of diatomic gas ($\gamma = 1.4$) at a pressure of 2 atmospheres is compressed adiabatically so that its temperature rises from 27°C to 927°C . The pressure of the gas in the final state is :-

- (1) 8 atm (2) 28 atm
 (3) 68.7 atm (4) 256 atm

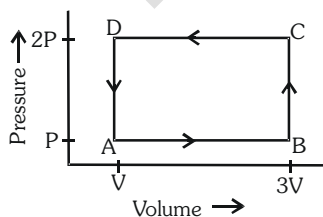
22. When 1 kg of ice at 0°C melts to water at 0°C , the resulting change in its entropy, taking latent heat of ice to be 80 cal/g , is -
- (1) 273 cal/K (2) $8 \times 10^4 \text{ cal/K}$
 (3) 80 cal/K (4) 293 cal/K

AIPMT (Pre) 2012

23. If the radius of a star is R and it acts as a black body, what would be the temperature of the star, in which the rate of energy production is Q ?
- (1) $(4\pi R^2 Q / \sigma)^{1/4}$ (2) $(Q / 4\pi R^2 \sigma)^{1/4}$
 (3) $Q / 4\pi R^2 \sigma$ (4) $(Q / 4\pi R^2 \sigma)^{-1/2}$
 (σ stands for Stefan's constant.)
24. One mole of an ideal gas goes from an initial state A to final state B via two processes. It firstly undergoes isothermal expansion from volume V to $3V$ and then its volume is reduced from $3V$ to V at constant pressure. The correct P-V diagram representing the two processes is :-

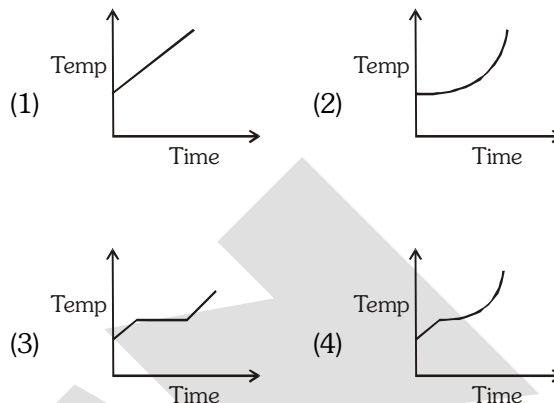


25. A thermodynamic system is taken through the cycle ABCD as shown in figure. Heat rejected by the gas during the cycle is :-



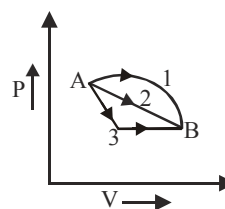
- (1) $\frac{1}{2} PV$ (2) PV (3) $2PV$ (4) $4PV$

26. Liquid oxygen at 50 K is heated to 300 K at constant pressure of 1 atm . The rate of heating is constant. Which one of the following graphs represents the variation of temperature with time ?



AIPMT (Mains) 2012

27. A slab of stone of area 0.36 m^2 and thickness 0.1 m is exposed on the lower surface to steam at 100°C . A block of ice at 0°C rests on the upper surface of the slab. In one hour 4.8 kg of ice is melted. The thermal conductivity of slab is :
 (Given latent heat of fusion of ice $3.36 \times 10^5 \text{ J kg}^{-1}$)
 (1) $2.05 \text{ J/m/s/}^\circ\text{C}$ (2) $1.02 \text{ J/m/s/}^\circ\text{C}$
 (3) $1.24 \text{ J/m/s/}^\circ\text{C}$ (4) $1.29 \text{ J/m/s/}^\circ\text{C}$
28. An ideal gas goes from state A to state B via three different processes as indicated in the P-V diagram



If Q_1 , Q_2 , Q_3 indicate the heat absorbed by the gas along the three processes and ΔU_1 , ΔU_2 , ΔU_3 indicate the change in internal energy along the three processes respectively, then :-

- (1) $Q_1 = Q_2 = Q_3$ and $\Delta U_1 > \Delta U_2 > \Delta U_3$
 (2) $Q_3 > Q_2 > Q_1$ and $\Delta U_1 > \Delta U_2 > \Delta U_3$
 (3) $Q_1 > Q_2 > Q_3$ and $\Delta U_1 = \Delta U_2 = \Delta U_3$
 (4) $Q_3 > Q_2 > Q_1$ and $\Delta U_1 = \Delta U_2 = \Delta U_3$

AIIMS 2012

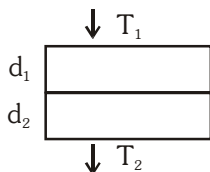
29. Two conductors having thickness d_1 and d_2 , thermal conductivity k_1 and k_2 are placed one above the another. Find the equivalent thermal conductance :-

(1) $\frac{(d_1 + d_2)(k_1 d_2 + k_2 d_1)}{2(k_1 + k_2)}$

(2) $\frac{(d_1 - d_2)(k_1 d_2 + k_2 d_1)}{2(k_1 + k_2)}$

(3) $\frac{k_1 k_2 (d_1 + d_2)}{d_1 k_2 + d_2 k_1}$

(4) None of these



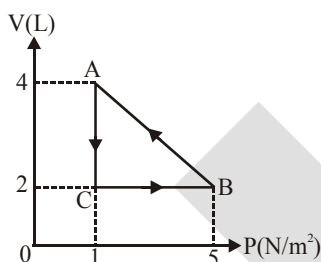
30. Degree of freedom for polyatomic gas :-

(1) ≥ 3 (2) ≥ 4 (3) ≥ 5 (4) ≥ 7

31. Conversion of water to steam is accompanied by which process :-

(1) Adiabatic (2) Isothermal
(3) Isochoric (4) Cyclic

32. Calculate the work done for $B \rightarrow A$, :-



(1) 6×10^{-3} J (2) 12×10^{-3} J
(3) 3×10^{-3} J (4) 4×10^{-3} J

33. What is the slope for an isothermal process in PV indicator diagram :-

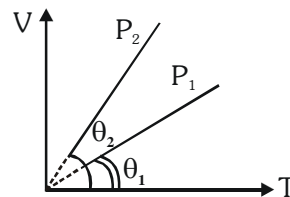
(1) $\frac{P}{V}$ (2) $-\frac{P}{V}$ (3) Zero (4) ∞

NEET-UG 2013

34. A piece of iron is heated in a flame. It first becomes dull red then becomes reddish yellow and finally turns to white hot. The correct explanation for the above observation is possible by using :-

(1) Newton's Law of cooling
(2) Stefan's Law
(3) Wein's displacement Law
(4) Kirchoff's Law

35. In the given ($V - T$) diagram, what is the relation between pressure P_1 and P_2 ?



(1) Cannot be predicted
(2) $P_2 = P_1$
(3) $P_2 > P_1$
(4) $P_2 < P_1$

36. The amount of heat energy required to raise the temperature of 1 g of Helium at constant volume, from T_1 K to T_2 K is :-

(1) $\frac{3}{4} N_a k_B \left(\frac{T_2}{T_1} \right)$

(2) $\frac{3}{8} N_a k_B (T_2 - T_1)$

(3) $\frac{3}{2} N_a k_B (T_2 - T_1)$

(4) $\frac{3}{4} N_a k_B (T_2 - T_1)$

37. The molar specific heats of an ideal gas at constant pressure and volume are denoted by C_p and C_v respectively. If $\gamma = \frac{C_p}{C_v}$ and R is the universal gas constant, then C_v is equal to :

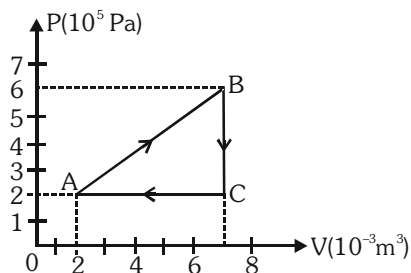
(1) γR (2) $\frac{1 + \gamma}{1 - \gamma}$

(3) $\frac{R}{(\gamma - 1)}$ (4) $\frac{(\gamma - 1)}{R}$

38. During an adiabatic process, the pressure of a gas is found to be proportional to the cube of its temperature. The ratio of $\frac{C_p}{C_v}$ for the gas is :-

(1) $\frac{3}{2}$ (2) $\frac{4}{3}$ (3) 2 (4) $\frac{5}{3}$

39. A gas is taken through the cycle $A \rightarrow B \rightarrow C \rightarrow A$, as shown, What is the net work done by the gas ?



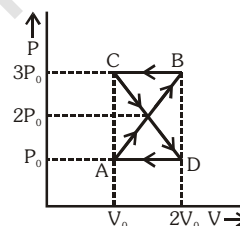
- (1) -2000 J (2) 2000 J
(3) 1000 J (4) Zero

AIIMS 2013

40. In a closed cylinder there are 60 g Ne and 64 g O_2 . If pressure of mixture of gases in cylinder is 30 bar then partial pressure of O_2 in this cylinder will be :-
(1) 30 bar (2) 20 bar
(3) 15 bar (4) 12 bar
41. A gas mixture contain one mole He gas and one mole O_2 gas. Find the ratio of specific heat at constant pressure to that at constant volume of the gaseous mixture :-
(1) 2 (2) 1.5
(3) 2.5 (4) 4
42. One mole of oxygen of volume 1 litre at 4 atm pressure to attain 1 atm pressure by result of isothermal expansion. Find work done by the gas.
(1) ~155 J (2) ~206 J
(3) ~355 J (4) ~562 J
43. Which of the following is a state function :-
(1) Work done in cyclic process
(2) Work done in isothermal process
(3) Heat at constant pressure
(4) Heat at constant volume
44. A refrigerator transfer 180 joule of energy in one second from temperature -3°C to 27°C . Calculate the average power consumed, assuming no energy losses in the process.
(1) 18 W (2) 54 W
(3) 20 W (4) 120 W

AIPMT 2014

45. Steam at 100°C is passed into 20 g of water at 10°C . When water acquires a temperature of 80°C , the mass of water present will be :
[Take specific heat of water = $1 \text{ cal g}^{-1} ^\circ\text{C}^{-1}$ and latent heat of steam = 540 cal g^{-1}]
(1) 24 g (2) 31.5 g
(3) 42.5 g (4) 22.5 g
46. Certain quantity of water cools from 70°C to 60°C in the first 5 minutes and to 54°C in the next 5 minutes. The temperature of the surroundings is:-
(1) 45°C (2) 20°C (3) 42°C (4) 10°C
47. The mean free path of molecules of a gas, (radius 'r') is inversely proportional to :-
(1) r^3 (2) r^2 (3) r (4) \sqrt{r}
48. A monoatomic gas at a pressure P, having a volume V expands isothermally to volume 2V and then adiabatically to volume 16V. The final pressure of the gas is : (take $\gamma = \frac{5}{3}$)
(1) 64P (2) 32P (3) $\frac{P}{64}$ (4) 16P
49. A thermodynamic system undergoes cyclic process ABCDA as shown in fig. The work done by the system in the cycle is :-



- (1) $P_0 V_0$ (2) $2P_0 V_0$
(3) $\frac{P_0 V_0}{2}$ (4) Zero

AIIMS 2014

50. Two stars A and B of surface area S_a and S_b & temperature T_a and T_b glow red and blue respectively. Choose the correct option.
(1) $T_a > T_b$ (2) $T_a < T_b$
(3) $T_a S_a = T_b S_b$ (4) $T_a S_b = T_b S_a$

51. A gas is equilibrium at T kelvin. If mass of one molecule is m and its component of velocity in x direction is v_x . Then mean of its v_x^2 is

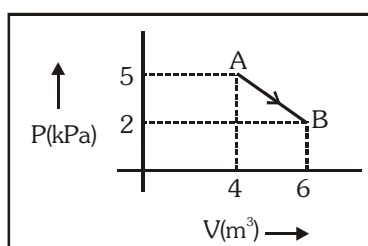
(1) $\frac{3kT}{m}$ (2) $\frac{2kT}{m}$ (3) $\frac{kT}{m}$ (4) zero

52. A diatomic gas undergoes adiabatic compression and its volume reduces to half of initial volume then its final pressure would be if initial pressure of gas is P .

(1) $2^{1.4}P$ (2) $P/2$
(3) $2P$ (4) $3.07 P$

AIPMT 2015

53. One mole of an ideal diatomic gas undergoes a transition from A to B along a path AB as shown in the figure,



The change in internal energy of the gas during the transition is :

(1) -20 kJ (2) 20 J (3) -12 kJ (4) 20 kJ

54. On observing light from three different stars P, Q and R, it was found that intensity of violet color is maximum in the spectrum of P, the intensity of green colour is maximum in the spectrum of R and the intensity of red colour is maximum in the spectrum of Q. If T_P , T_Q and T_R are the respective absolute temperatures of P, Q and R, then it can be concluded from the above observation that :

(1) $T_P > T_R > T_Q$ (2) $T_P < T_R < T_Q$
(3) $T_P < T_Q < T_R$ (4) $T_P > T_Q > T_R$

55. A Carnot engine, having an efficiency of $\eta = \frac{1}{10}$ as heat engine, is used as a refrigerator. If the work done on the system is 10 J , the amount of energy absorbed from the reservoir at lower temperature is :-

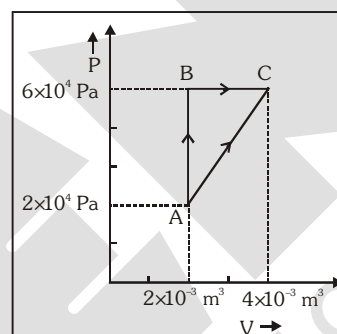
(1) 99 J (2) 90 J (3) 1 J (4) 100 J

56. The ratio of the specific heats $\frac{C_P}{C_V} = \gamma$ in terms of degrees of freedom (n) is given by :

(1) $\left(1 + \frac{n}{3}\right)$ (2) $\left(1 + \frac{2}{n}\right)$

(3) $\left(1 + \frac{n}{2}\right)$ (4) $\left(1 + \frac{1}{n}\right)$

57. Figure below shows two paths that may be taken by a gas to go from a state A to a state C.



In process AB, 400 J of heat is added to the system and in process BC, 100 J of heat is added to the system. The heat absorbed by the system in the process AC will be :

(1) 500 J (2) 460 J
(3) 300 J (4) 380 J

58. The two ends of a metal rod are maintained at temperatures 100°C and 110°C . The rate of heat flow in the rod is found to be 4.0 J/s . If the ends are maintained at temperatures 200°C and 210°C , the rate of heat flow will be :

(1) 16.8 J/s (2) 8.0 J/s
(3) 4.0 J/s (4) 44.0 J/s

Re-AIPMT 2015

59. Two vessels separately contain two ideal gases A and B at the same temperature, the pressure of A being twice that of B. Under such conditions, the density of A is found to be 1.5 times the density of B. The ratio of molecular weight of A and B is:

(1) $\frac{1}{2}$ (2) $\frac{2}{3}$ (3) $\frac{3}{4}$ (4) 2

72. A gas is compressed isothermally to half its initial volume. The same gas is compressed separately through an adiabatic process until its volume is again reduced to half. Then :-

- (1) Compressing the gas isothermally will require more work to be done.
- (2) Compressing the gas through adiabatic process will require more work to be done.
- (3) Compressing the gas isothermally or adiabatically will require the same amount of work.
- (4) Which of the case (whether compression through isothermal or through adiabatic process) requires more work will depend upon the atomicity of the gas.

73. A piece of ice falls from a height h so that it melts completely. Only one-quarter of the heat produced is absorbed by the ice and all energy of ice gets converted into heat during its fall. The value of h is :

[Latent heat of ice is 3.4×10^5 J/kg and

$g = 10$ N/kg]

- (1) 34 km
- (2) 544 km
- (3) 136 km
- (4) 68 km

NEET-II 2016

74. Two identical bodies are made of a material for which the heat capacity increases with temperature. One of these is at 100°C , while the other one is at 0°C . If the two bodies are brought into contact, then, assuming no heat loss, the final common temperature is :-

- (1) less than 50°C but greater than 0°C
- (2) 0°C
- (3) 50°C
- (4) more than 50°C

75. A body cools from a temperature $3T$ to $2T$ in 10 minutes. The room temperature is T . Assume that Newton's law of cooling is applicable. The temperature of the body at the end of next 10 minutes will be :-

- (1) $\frac{4}{3}T$
- (2) T
- (3) $\frac{7}{4}T$
- (4) $\frac{3}{2}T$

76. One mole of an ideal monatomic gas undergoes a process described by the equation $PV^3 = \text{constant}$. The heat capacity of the gas during this process is

- (1) $2R$
- (2) R
- (3) $\frac{3}{2}R$
- (4) $\frac{5}{2}R$

77. The temperature inside a refrigerator is $t_2^\circ\text{C}$ and the room temperature is $t_1^\circ\text{C}$. The amount of heat delivered to the room for each joule of electrical energy consumed ideally will be :-

- (1) $\frac{t_2 + 273}{t_1 - t_2}$
- (2) $\frac{t_1 + t_2}{t_1 + 273}$
- (3) $\frac{t_1}{t_1 - t_2}$
- (4) $\frac{t_1 + 273}{t_1 - t_2}$

78. A given sample of an ideal gas occupies a volume V at a pressure P and absolute temperature T . The mass of each molecule of the gas is m . Which of the following gives the density of the gas ?

- (1) $P/(kTV)$
- (2) mkT
- (3) $P/(kT)$
- (4) $Pm/(kT)$

AIIMS 2016

79. In adiabatic process, volume of monoatomic gas increases by 6% then find percentage change in temperature.

- (1) -2%
- (2) -4%
- (3) 2%
- (4) 4%

92. At constant pressure how much fraction of heat supplied to gas is converted into mechanical work?

- (1) $\frac{\gamma-1}{\gamma}$ (2) $\frac{\gamma}{\gamma-1}$
(3) $\gamma-1$ (4) $\frac{\gamma}{\gamma+1}$

NEET(UG)-2018

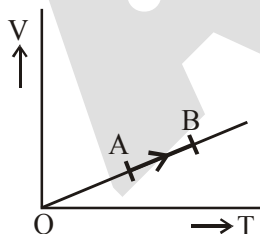
93. A sample of 0.1 g of water at 100°C and normal pressure ($1.013 \times 10^5 \text{ Nm}^{-2}$) requires 54 cal of heat energy to convert to steam at 100°C. If the volume of the steam produced is 167.1 cc, the change in internal energy of the sample, is :-

- (1) 104.3 J (2) 208.7 J
(3) 42.2 J (4) 84.5 J

94. The power radiated by a black body is P and it radiates maximum energy at wavelength λ_0 . If the temperature of the black body is now changed so that it radiates maximum energy at wavelength $\frac{3}{4}\lambda_0$, the power radiated by it becomes nP. The value of n is :-

- (1) $\frac{3}{4}$ (2) $\frac{4}{3}$
(3) $\frac{256}{81}$ (4) $\frac{81}{256}$

95. The volume (V) of a monatomic gas varies with its temperature (T), as shown in the graph. The ratio of work done by the gas, to the heat absorbed by it, when it undergoes a change from state A to state B, is :-



- (1) $\frac{2}{5}$ (2) $\frac{2}{3}$ (3) $\frac{1}{3}$ (4) $\frac{2}{7}$

96. The efficiency of an ideal heat engine working between the freezing point and boiling point of water, is :-

- (1) 26.8% (2) 20% (3) 6.25% (4) 12.5%

97. At what temperature will the rms speed of oxygen molecules become just sufficient for escaping from the Earth's atmosphere ?

(Given :

Mass of oxygen molecule (m) = $2.76 \times 10^{-26} \text{ kg}$
Boltzmann's constant $k_B = 1.38 \times 10^{-23} \text{ J K}^{-1}$:-

- (1) $2.508 \times 10^4 \text{ K}$ (2) $8.360 \times 10^4 \text{ K}$
(3) $5.016 \times 10^4 \text{ K}$ (4) $1.254 \times 10^4 \text{ K}$

AIIMS 2018

98. Which of the following properties of a thermodynamic system are extensive ?

- (1) V and T (2) P and T
(3) V and E (4) E and C_p

99. 1 cm³ of water at 100°C and 1 atm pressure is heated to convert it to 1521 cm³ vapour. Find increase in internal energy if latent heat of vaporization of water = 540 cal/g.

- (1) 3.26 kJ (2) 2.12 kJ
(3) 4.18 kJ (4) 1.05 kJ

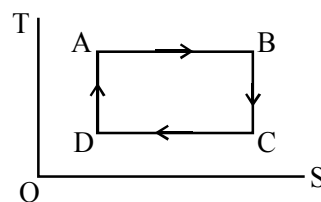
100. Which of the following is intensive property of a thermodynamic system ?

- (1) P and T (2) V and T
(3) V and P (4) V and E

101. Work done in expanding one mole of an ideal gas isothermally from 2L to 4L is equal to the work done in expanding three moles of an ideal gas from 2L to x L at same temperature. Find x.

- (1) $8^{1/3}$ (2) $4^{2/3}$ (3) 2 (4) 16

102. Consider the given temperature entropy diagram. Which of the following processes represent increase in volume ?

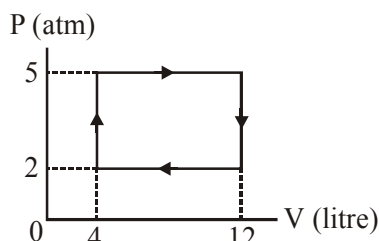


- (1) DA and AB (2) AB and BC
(3) BC and CD (4) CD and DA

103. Nitrogen gas is contained in a cylinder at 8 atm pressure and 300 K. The density of nitrogen in the container is nearly :-

- (1) 7.9 kg/m³ (2) 9 kg/m³
(3) 1.3 kg/m³ (4) 12 kg/m³

- 104.** An ideal gas undergoes a cyclic process as shown in diagram. The net work done by the gas in the cycle is :-



- (1) 12 litre-atm (2) 24 J
(3) 24 litre-atm (4) -24 J
- 105.** A thermodynamic system does 60 J of useful work. If its internal energy decreases by 50J, then how much heat has been supplied to the system ?
(1) 10 J (2) 110 J (3) -10 J (4) -60 J
- 106.** In an adiabatic expansion of 0.5 mole of an ideal diatomic gas, temperature changes from 32°C to 20°C. The change in internal energy of the gas is
(1) 124.65 cal. (2) -124.65 J
(3) 72.32 J (4) 248.3 J

- 107.** A refrigerator having coefficient of performance 10 absorbs 500 J per cycle. The amount of heat rejected to the surrounding per cycle is :-

- (1) 50 J (2) 550 J
(3) 450 J (4) 500 J

- 108.** Two gases A and B have same temperature. Density of A is twice the density of B but molar mass of B is twice as that of A. The ratio of pressures of gases A and B is :-

- (1) 1 : 1 (2) 2 : 1
(3) 4 : 1 (4) 1 : 4

- 109.** The change in entropy of 1 kg of nitrogen gas in an isobaric process when its temperature changes from 400 K to 800 K is :-

- (1) 720 J/K (2) 420 J/K
(3) 309 J/K (4) 720 cal/°C

EXERCISE-II (Previous Years Questions)

ANSWER KEY

Que.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15
Ans.	2	2	1	2	2	3	2	4	2	1	1	4	3	1	3
Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	2	2	2	3	4	4	4	2	2	3	3	3	3	3	3
Que.	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45
Ans.	2	1	2	3	4	2	3	1	3	4	2	4	4	3	4
Que.	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60
Ans.	1	2	3	4	2	3	1	1	1	2	2	2	3	3	2
Que.	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75
Ans.	2	2	3	2	3	1	3	3	4	4	2	2	3	4	4
Que.	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90
Ans.	2	4	4	2	1	3	4	2	4	1	3	1	3	1	4
Que.	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105
Ans.	3	1	2	3	1	1	2	3	2	1	2	2	2	3	1
Que.	106	107	108	109											
Ans.	2	2	3	1											

EXERCISE-III (Analytical Questions)

Check Your Understanding

1. Two rods one of aluminium of length l_1 having coefficient of linear expansion α_a , and other steel of length l_2 having coefficient of linear expansion α_s are joined end to end. The expansion in both the rods is same on variation of temperature. Then the

value of $\frac{l_1}{l_1 + l_2}$ is

- (1) $\frac{\alpha_s}{\alpha_a + \alpha_s}$ (2) $\frac{\alpha_s}{\alpha_a - \alpha_s}$
(3) $\frac{\alpha_a + \alpha_s}{\alpha_s}$ (4) None of these

2. An ideal gas is expanding such that $PT^2 = \text{constant}$. The coefficient of volume expansion of the gas is :

- (1) $\frac{1}{T}$ (2) $\frac{2}{T}$
(3) $\frac{3}{T}$ (4) $\frac{4}{T}$

3. Steam at 100°C is added slowly to 1400 g of water at 16°C until the temperature of water is raised to 80°C . The mass of steam required to do this is ($L_v = 540 \text{ cal/g}$) :

- (1) 160 g (2) 125 g
(3) 250 g (4) 320 g

4. 50 g of ice at 0°C is mixed with 50 g of water at 100°C . The final temperature of mixture is :-

- (1) 0°C
(2) Between 0°C to 20°C
(3) 20°C
(4) Above 20°C

5. In an insulated vessel, 0.05 kg steam at 373 K and 0.45 kg of ice at 253 K are mixed. Then, find the final temperature of the mixture.

Given, $L_{\text{fusion}} = 80 \text{ cal/g} = 336 \text{ J/g}$, $L_{\text{vaporization}} = 540 \text{ cal/g} = 2268 \text{ J/g}$, $S_{\text{ice}} = 2100 \text{ J/kg K} = 0.5 \text{ cal/gK}$ and $S_{\text{water}} = 4200 \text{ J/kg K} = 1 \text{ cal/gK}$

- (1) 273 K (2) 373 K
(3) 100 K (4) 0 K

6. Spheres P and Q are uniformly constructed from the same material which is a good conductor of heat and the radius of Q is thrice of radius of P. The rate of fall of temperature of P is x times that of Q when both are at the same surface temperature. The value of x is :

- (1) 1/4 (2) 1/3 (3) 3 (4) 4

7. A sphere, a cube and a thin circular plate all made of same substance and all have same mass. These are heated to 200°C and then placed in a room, then the :

- (1) Temperature of sphere drops to room temperature at last.
(2) Temperature of cube drops to room temperature at last
(3) Temperature of thin circular plate drops to room temperature at last
(4) Temperature of all the three drops to room temperature at the same time

8. The power radiated by a black body is P and it radiates maximum energy around the wavelength λ_0 . If the temperature of the black body is now changed so that it radiates maximum energy around wavelength $3/4\lambda_0$, the power radiated by it will increase by a factor of

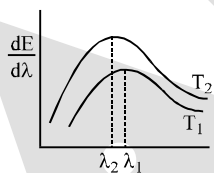
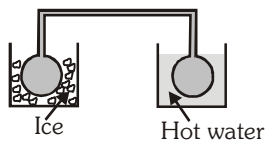
- (1) 4/3 (2) 16/9
(3) 64/27 (4) 256/81

9. A slab consists of two parallel layers of copper and brass of the same thickness and having thermal conductivities in the ratio 1 : 4. If the free face of brass is at 100°C and that of copper at 0°C , the temperature of interface is -

- (1) 80°C (2) 20°C (3) 60°C (4) 40°C

10. Out of the metal balls of same diameter one is solid and other is hollow. Both are heated to the same temperature at 300°C and then allowed to cool in the same surroundings then rate of loss of heat will be:

- (1) More for hollow sphere
(2) More for solid sphere
(3) Same for both
(4) None of the above

- 11.** For a black body at temperature 727°C , its radiating power is 60 watt and temperature of surrounding is 227°C . If temperature of black body is changed to 1227°C then its radiating power will be :-
 (1) 304 W (2) 320 W
 (3) 240 W (4) 120 W
- 12.** A body takes T minutes to cool from 62°C to 61°C when the surrounding temperature is 30°C . The time taken by the body to cool from 46°C to 45.5°C is :-
 (1) Greater than T minutes
 (2) Equal to T minutes
 (3) Less than T minutes
 (4) Equal to $T/2$ minutes
- 13.** Three discs A, B, and C having radii 2 m, 4 m and 6 m respectively are coated with carbon black on their outer surfaces. The wavelengths corresponding to maximum intensity are 300 nm, 400 nm and 500 nm respectively. The power radiated by them are Q_A , Q_B and Q_C respectively then
 (1) Q_A is maximum (2) Q_B is maximum
 (3) Q_C is maximum (4) $Q_A = Q_B = Q_C$
- 14.** The spectral emissive power E_{λ} for a body at temperature T_1 is plotted against the wavelength and area under the curve is found to be A . At a different temperature T_2 , the area is found to be $9A$. Then $\lambda_1/\lambda_2 =$
- 
- (1) 3 (2) $1/3$ (3) $1/\sqrt{3}$ (4) $\sqrt{3}$
- 15.** Consider a gas with density ρ and \bar{c} as the root mean square velocity of its molecules contained in a volume. If the system moves as whole with velocity v , then the pressure exerted by the gas is
 (1) $\frac{1}{3} \rho (\bar{c})^2$ (2) $\frac{1}{3} \rho (\bar{c} + v)^2$
 (3) $\frac{1}{3} \rho (\bar{c} - v)^2$ (4) $\frac{1}{3} \rho (\bar{c}^2 - v^2)$
- 16.** Air is filled at 60°C in a vessel of open mouth. The vessel is heated to a temperature T so that $1/4^{\text{th}}$ part of air escapes. The value of T is :
 (1) 80°C (2) 444°C
 (3) 333°C (4) 171°C
- 17.** Root mean square velocity for a certain di-atomic gas at room temperature 27°C is found to be 1930 m/sec. The gas is -
 (1) H_2 (2) O_2 (3) F_2 (4) Cl_2
- 18.** The equation of state of a gas is given by $\left(P + \frac{aT^2}{V}\right)V^c = (RT + b)$, where a , b , c and R are constants. This isotherms can be represented by $P = AV^m - BV^n$, where A and B depend only on temperature and
 (1) $m = -c$ and $n = -1$
 (2) $m = c$ and $n = 1$
 (3) $m = -c$ and $n = 1$
 (4) $m = c$ and $n = -1$
- 19.** An ideal gas expands in such a way that $PV^2 = \text{constant}$ throughout the process.
 (1) The graph of the process of T - V diagram is a parabola.
 (2) The graph of the process of T - V diagram is a straight line.
 (3) Such an expansion is possible only with heating.
 (4) Such an expansion is possible only with cooling.
- 20.** Two identical glass bulbs are interconnected by a thin glass tube at 0°C . A gas is filled at N.T.P. in these bulb is placed in ice and another bulb is placed in hot bath, then the pressure of the gas becomes 1.5 times. The temperature of hot bath will be :-
- 
- (1) 100°C (2) 182°C
 (3) 256°C (4) 546°C

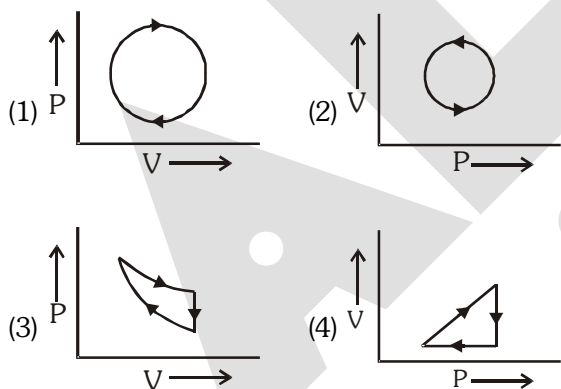
21. A gas has volume V and pressure P . The total translational kinetic energy of all the molecules of the gas is:-

- (1) $\frac{3}{2}PV$ only if the gas is monoatomic.
- (2) $\frac{3}{2}PV$ only if the gas is diatomic.
- (3) $>\frac{3}{2}PV$ if the gas is diatomic.
- (4) $\frac{3}{2}PV$ in all cases.

22. N molecules of an ideal gas at temperature T_1 and pressure P_1 are contained in a closed box. If the molecules in the box gets doubled, Keeping total kinetic energy same. If new pressure is p_2 and temperature is T_2 , Then :

- (1) $P_2 = P, T_2 = T_1$
- (2) $P_2 = P_1, T_2 = T_1 / 2$
- (3) $P_2 = 2 P_1, T_2 = T_1$
- (4) $P_2 = 2P_1, T_2 = T_1 / 2$

23. The following are the P - V diagrams for cyclic processes for a gas. In which of these processes heat is released by the gas ?



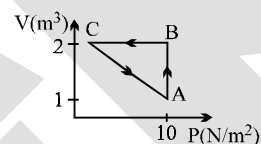
24. A gas mixture contain 1 g H_2 and 1 g He if temperature of gas mixture is increased from $0^\circ C$ to $100^\circ C$ at isobaric process. Then find given heat of gas mixture

- $[\gamma_{He} = 5/3, \gamma_{H_2} = 7/5, R = 2 \text{ cal/mol-K}]$
- (1) 124 cal
 - (2) 327 cal
 - (3) 218 cal
 - (4) 475 cal

25. An ideal gas expands isothermally from a volume V_1 to V_2 and then compressed to original volume V_1 adiabatically. Initial pressure is P_1 and final pressure is P_3 . The total work done is W . Then

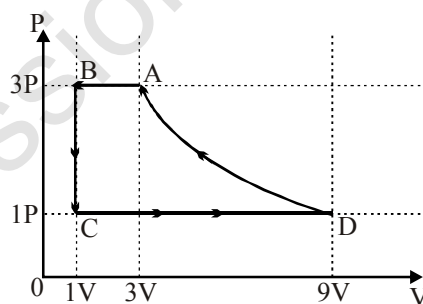
- (1) $P_3 > P_1, W > 0$
- (2) $P_3 < P_1, W < 0$
- (3) $P_3 > P_1, W < 0$
- (4) $P_3 = P_1, W = 0$

26. An ideal gas is taken through the cycle $A \rightarrow B \rightarrow C \rightarrow A$, as shown in the figure. If the net heat supplied to the gas in the cycle is 5J, the work done by the gas in the process $C \rightarrow A$ is



- (1) -5J
- (2) -10 J
- (3) -15 J
- (4) -20 J

27. One mole of a monatomic ideal gas is taken through a cycle ABCDA as shown in the P - V diagram. Column-II gives the characteristics involved in the cycle. Match them with each of the processes given in Column-I.

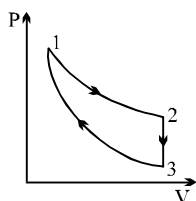


Column-I

Column-II

- (A) Process $A \rightarrow B$ (p) Internal energy decreases
 (B) Process $B \rightarrow C$ (q) Internal energy increases
 (C) Process $C \rightarrow D$ (r) Heat is lost
 (D) Process $D \rightarrow A$ (s) Heat is gained
 (t) Work is done on the gas
- (1) $A \rightarrow p, r, t$; $B \rightarrow p, r$; $C \rightarrow q, s$; $D \rightarrow r, t$
 - (2) $A \rightarrow r, t$; $B \rightarrow q, r$; $C \rightarrow q, s$; $D \rightarrow r, s$
 - (3) $A \rightarrow p, r, t$; $B \rightarrow p, q, r$; $C \rightarrow s$; $D \rightarrow r, t$
 - (4) $A \rightarrow p, r, t$; $B \rightarrow p, r$; $C \rightarrow q$; $D \rightarrow r, s$

28. Three processes form a thermodynamic cycle as shown on P-V diagram for an ideal gas. Process $1 \rightarrow 2$ takes place at constant temperature (300K). Process $2 \rightarrow 3$ takes place at constant volume. During this process 40J of heat leaves the system. Process $3 \rightarrow 1$ is adiabatic and temperature T_3 is 275K. Work done by the gas during the process $3 \rightarrow 1$ is

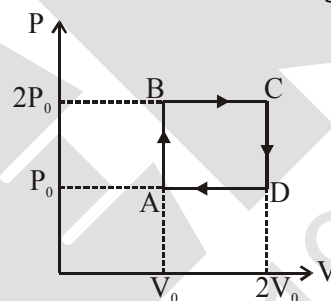


- (1) -40J (2) -20J (3) +40J (4) +20J
29. The molar specific heat under constant pressure of oxygen is $C_p = 7.03 \text{ cal/mol K}$. The quantity of heat required to raise the temperature from 10°C to 20°C of 5 moles of oxygen under constant volume will approximately be :-
- (1) 25 cal (2) 50 cal
(3) 250 cal (4) 500 cal
30. If the ratio of specific heat of a gas at constant pressure to that at constant volume is γ , the change in internal energy of a mass of gas, when the volume changes from V to $2V$ at constant pressure P , is :-
- (1) $R/(\gamma - 1)$ (2) PV
(3) $PV/(\gamma - 1)$ (4) $\gamma PV/(\gamma - 1)$

31. The molar heat capacity in a process of a diatomic gas, if it does a work of $\frac{Q}{4}$ when a heat of Q is supplied to it, is :-

- (1) $\frac{2}{5}R$ (2) $\frac{5}{2}R$
(3) $\frac{10}{3}R$ (4) $\frac{6}{7}R$

32. The above P-V diagram represents the thermodynamic cycle of an engine, operating with an ideal monoatomic gas. The amount of heat, extracted from the source in a single cycle is :



- (1) P_0V_0 (2) $\left(\frac{13}{2}\right)P_0V_0$
(3) $\left(\frac{11}{2}\right)P_0V_0$ (4) $4P_0V_0$
33. In the above question efficiency of cycle ABCDA is nearly :
- (1) 12.5% (2) 15.4%
(3) 9.1% (4) 10.5%

EXERCISE-III (Analytical Questions)

ANSWER KEY

Que.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15
Ans.	1	3	1	2	1	3	1	4	1	3	2	2	2	4	1
Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	4	1	1	4	4	4	2	4	4	3	1	1	1	3	3
Que.	31	32	33												
Ans.	3	2	2												

EXERCISE-IV (Assertion & Reason)

Target AIIMS

Directions for Assertion & Reason questions

These questions consist of two statements each, printed as Assertion and Reason. While answering these Questions you are required to choose any one of the following four responses.

- (A) If both Assertion & Reason are True & the Reason is a correct explanation of the Assertion.
 (B) If both Assertion & Reason are True but Reason is not a correct explanation of the Assertion.
 (C) If Assertion is True but the Reason is False.
 (D) If both Assertion & Reason are false.

1. **Assertion :** Liquids usually expand more than solids.

Reason : The intermolecular forces in liquids are weaker than in solids.

- (1) A (2) B (3) C (4) D

2. **Assertion :** Rubber contract on heating.

Reason : In rubber as temperature increases, the amplitude of transverse vibrations increases more than the amplitude of longitudinal vibrations.

- (1) A (2) B (3) C (4) D

3. **Assertion :** A temperature change which increases the length of a steel rod by 1% will increase its volume by 3%.

Reason : The coefficient of volume expansion is nearly three times the coefficient of linear expansion.

- (1) A (2) B (3) C (4) D

4. **Assertion :** The melting point of ice decreases with increases of pressure.

Reason : Ice contracts on melting. [AIIMS 2004]

- (1) A (2) B (3) C (4) D

5. **Assertion :** Water kept in an open vessel will quickly evaporate on the surface of the moon.

Reason : There is lack of atmospheric pressure at surface of moon.

- (1) A (2) B (3) C (4) D

6. **Assertion :** In a pressure cooker the water is brought to boil. The cooker is then removed from the stove. Now, on removing the lid of the pressure cooker, the water starts boiling again. [AIIMS 2004]

Reason : When we add small impurities in water, the boiling point of water will decrease.

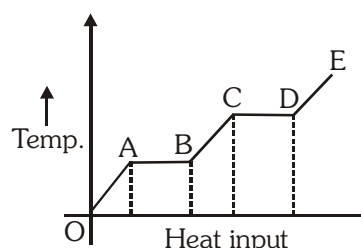
- (1) A (2) B (3) C (4) D

7. **Assertion :** In pressure-temperature phase diagram of water, the slope of the melting curve is found to be negative. [AIIMS 2005]

Reason : Ice contracts on melting of water.

- (1) A (2) B (3) C (4) D

8. **Assertion :** A solid material is supplied heat at constant rate. The temperature of the material is changing with the heat input as shown in Figure.



If $CD = 2AB$ then latent heat of vaporisation of substance is double that of fusion.

Reason : Latent heat of fusion $\propto AB$ and Latent heat of vaporisation $\propto CD$

- (1) A (2) B (3) C (4) D

- 9. Assertion :** It is hotter over the top of a fire than at the same distance on the sides. [AIIMS 2003]
Reason : Air surrounding the fire transfer more heat upwards.
 (1) A (2) B (3) C (4) D
- 10. Assertion :** Coolant coils are fitted at the top of a refrigerator, for formation of convection current.
Reason : Air becomes denser on cooling.
 (1) A (2) B (3) C (4) D
- 11. Assertion :** Perspiration from human body helps in cooling the body. [AIIMS 2006]
Reason : A thin layer of water on the skin enhances its emissivity.
 (1) A (2) B (3) C (4) D
- 12. Assertion :** A sphere, a cube and a thin circular plate made of same material and of same mass are initially heated to 200°C , the plate will cool at fastest rate.
Reason : Rate of cooling = $\frac{eA\sigma}{ms} (T^4 - T_0^4) \propto \text{surface area}$ and surface area is maximum for plate.
 (1) A (2) B (3) C (4) D
- 13. Assertion :** Blue star is at high temperature than red star.
Reason : According to Wien's displacement law
 $\lambda_m = \frac{b}{T}$, where symbols have their usual meanings.
 (1) A (2) B (3) C (4) D
- 14. Assertion :** A body that is a good radiator is also a good absorber of radiation at a given wavelength.
Reason : According to Kirchhoff's law the absorptivity of a body is equal to its emissivity at a given wavelength. [AIIMS 2005]
 (1) A (2) B (3) C (4) D
- 15. Assertion :** For higher temperature, the peak emission wavelength of a black body shifts to lower wavelengths. [AIIMS 2005]
Reason : Peak emission wavelength of a black body is proportional to the fifth power of temperature.
 (1) A (2) B (3) C (4) D
- 16. Assertion :** While measuring the thermal conductivity of liquid experimentally, the upper layer is kept hot and the lower layer is kept cold.
Reason : This avoids heating of liquid by convection. [AIIMS 2007]
 (1) A (2) B (3) C (4) D
- 17. Assertion :** The amount of radiation from sun's surface varies as the fourth power of its absolute temperature.
Reason : The sun is black body. [AIIMS 1999]
 (1) A (2) B (3) C (4) D
- 18. Assertion :** If a gas container in motion is suddenly stopped, the temperature of the gas rises.
Reason : In given process, the kinetic energy of ordered mechanical motion is converted into the kinetic energy of random motion of gas molecules.
 (1) A (2) B (3) C (4) D
- 19. Assertion :** A material will have only one specific heat, if and only if its coefficient of thermal expansion is equal to zero.
Reason : An ideal gas has two specific heats (C_v and C_p) only.
 (1) A (2) B (3) C (4) D
- 20. Assertion :** The value of ΔQ is always zero in adiabatic process. [AIIMS 2010]
Reason : Adiabatic process is always a cyclic process.
 (1) A (2) B (3) C (4) D
- 21. Assertion :** First law of thermodynamics allows many processes which actually don't happen.
Reason : First law of thermodynamics must not be violated for any process to happen.
 (1) A (2) B (3) C (4) D

22. Assertion : Thermodynamics processes in nature are irreversible. [AIIMS 2004]

Reason : Dissipative effects cannot be eliminated.

- (1) A (2) B (3) C (4) D

23. Assertion : Reversible system are difficult to find in real world. [AIIMS 2005]

Reason : Most processes are dissipative in nature.

- (1) A (2) B (3) C (4) D

24. Assertion : Air quickly leaking out of a ballon becomes cooler. [AIIMS 2005]

Reason : The leaking air undergoes adiabatic expansion.

- (1) A (2) B (3) C (4) D

25. Assertion : When a glass of hot milk is placed in a room and allowed to cool, its entropy decreases.

Reason : Allowing hot object to cool does not violate the second law of thermodynamics.

[AIIMS 2006]

- (1) A (2) B (3) C (4) D

26. Assertion : The Carnot cycle is useful in understanding the performance of heat engines.

Reason : The Carnot cycle provides a way of determinig the maximum possible efficiency achievable with reservoirs of given temperatures.

[AIIMS 2006]

- (1) A (2) B (3) C (4) D

27. Assertion : In an isolated system, the entropy increases.

Reason : The processes in an isolated system are adiabatic. [AIIMS 2006]

- (1) A (2) B (3) C (4) D

28. Assertion : Two isothermal curves intersect each other at a certain point. [AIIMS 2008]

Reason : The isothermal change are done slowly, so the isothermal curves have very little slope.

- (1) A (2) B (3) C (4) D

29. Assertion : It is not possible for a system, unaided by an external agency to transfer heat from a body at lower temperature to another at a higher temperature. [AIIMS 1994]

Reason : It is not possible to violate the second law of thermodynamics.

- (1) A (2) B (3) C (4) D

30. Assertion : It is impossible for a ship to use the internal energy of sea water to operate its engine.

Reason : A refrigerator is a heat engine working in the reverse direction. [AIIMS 2009]

- (1) A (2) B (3) C (4) D

31. Assertion : When a bottle of cold carbonated drink is opened, a slight fog forms around the opening.

Reason : Adiabatic expansion of the gas causes lowering of temperature which start condensation of water vapours. [AIIMS 2003]

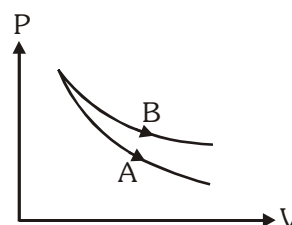
- (1) A (2) B (3) C (4) D

32. Assertion : In isothermal process heat energy supplied to system is converted completely into work.

Reason : According to first law of thermodynamics $Q = W + \Delta U$.

- (1) A (2) B (3) C (4) D

33. Assertion : The curve A and B in figure, show P-V graphs for an isothermal and an adiabatic process for an ideal gas. The isothermal process is represented by B.



Reason : On P-V graph, modulus of slope of the adiabatic curve is greater then the modulus of the slope of the isothermal curve.

- (1) A (2) B (3) C (4) D

- 34. Assertion :** During rapid pumping of air in tyres, air inside the tyre is hotter than atmospheric air.
Reason : Adiabatic process occurs at very high rate.
 (1) A (2) B (3) C (4) D
- 35. Assertion :** The root mean square and most probable speeds of the molecules in a gas are the same. [AIIMS 2006]
Reason : The maxwell's distribution for the speed of molecules in a gas is symmetrical.
 (1) A (2) B (3) C (4) D
- 36. Assertion :** The ratio $\frac{C_P}{C_V}$ for a monoatomic gas is more then for a diatomic gas. [AIIMS 1998]
Reason : The molecules of a monoatomic gas have more degree of freedom then those of a diatomic gas.
 (1) A (2) B (3) C (4) D
- 37. Assertion :** Air pressure in a car tyre increases during driving. [AIIMS 2007]
Reason : Absolute zero temperature is not zero energy temperaure.
 (1) A (2) B (3) C (4) D
- 38. Assertion :** The ratio $\frac{C_P}{C_V}$ is more for helium gas than for hydrogen gas.
Reason : Atomic mass of helium is more than that of hydrogen. [AIIMS 1996]
 (1) A (2) B (3) C (4) D
- 39. Assertion :** For an ideal gas $C_P > C_V$.
Reason : Work is done in expansion of the gas at constant pressure. [AIIMS 2011]
 (1) A (2) B (3) C (4) D
- 40. Assertion :** For an ideal gas, at constant temperature, the product of the pressure and volume is constant. [AIIMS 1998]
Reason : The mean square velocity of gas molecules is inversely proportional to its mass.
 (1) A (2) B (3) C (4) D
- 41. Assertion :** Water cannot be boiled inside a satellite by convection.
Reason : In weightlessness conditions, natural movement of heated fluid is not possible.
 (1) A (2) B (3) C (4) D
- 42. Assertion :-** The value of specific heat at constant pressure is more than that at constant volume.
Reason :- The energy required to raise temperature by a unit at constant pressure is greater because some amount of heat is used in doing work. [AIIMS 2015]
 (1) A (2) B (3) C (4) D
- 43. Assertion :-** A good absorber is a good emitter.
Reason :- For a substance, absorbance is equal to emissivity. [AIIMS 2015]
 (1) A (2) B (3) C (4) D
- 44. Assertion :-** Sum of two path functions is always a path function.
Reason :- Every state function can be produced by path functions. [AIIMS 2015]
 (1) A (2) B (3) C (4) D
- 45. Assertion :-** Specific heat of metals approaches zero as temperature approaches zero kelvin.
Reason :- As tempeature approaches zero kelvin, effectiveness of degree of freedom decreases. [AIIMS 2015]
 (1) A (2) B (3) C (4) D
- 46. Assertion :-** When we apply brakes to car its entropy doesn't increase. [AIIMS 2016]
Reason :- It is a reversible process.
 (1) A (2) B (3) C (4) D
- 47. Assertion :-** A fluid in vacuum expands such that it changes to gaseous state. Its temperature decreases. [AIIMS 2016]
Reason :- It uses its internal energy as heat of vaporization.
 (1) A (2) B (3) C (4) D

48. **Assertion :-** H_2 is found more as compared to N_2 in earth's upper atmosphere. [AIIMS 2016]

Reason :- N_2 is heavier than H_2 .

- (1) A (2) B (3) C (4) D

49. **Assertion :-** PV value of gas can increase on increasing the pressure. [AIIMS 2016]

Reason :- Pressure always increases on compression.

- (1) A (2) B (3) C (4) D

50. **Assertion :-** Heat is not a state thermodynamic function. [AIIMS 2016]

Reason :- Heat given in a system depends on process.

- (1) A (2) B (3) C (4) D

51. **Assertion :-** In an adiabatic process, temperature of system decrease when work is done by the system. [AIIMS 2016]

Reason :- A part of internal energy is utilized in doing work.

- (1) A (2) B (3) C (4) D

52. **Assertion :-** For an isolated system, mechanical energy is not conserved. [AIIMS 2017]

Reason :- For an isolated system internal energy is constant.

- (1) A (2) B (3) C (4) D

53. **Assertion :-** Adiabatic free expansion is not a reversible process. [AIIMS 2017]

Reason :- In adiabatic expansion internal energy does not change.

- (1) A (2) B (3) C (4) D

54. **Assertion :-** Generally, molecules which have maximum velocity are lesser in number.

Reason :- The molecules of the gas follow Maxwell's speed distribution. [AIIMS 2017]

- (1) A (2) B (3) C (4) D

55. **Assertion :-** Adiabatic free expansion of an ideal gas is irreversible. [AIIMS 2017]

Reason :- $PV = \text{constant}$ for reversible adiabatic expansion.

- (1) A (2) B (3) C (4) D

56. **Assertion :-** Thermal energy and heat energy are same. [AIIMS 2018]

Reason :- Both energies depend on randomness of molecules.

- (1) A (2) B (3) C (4) D

57. **Assertion :-** Temperature is constant in a Quasi-static isothermal process. [AIIMS 2018]

Reason :- No heat is exchanged with surrounding in a quasi-static isothermal process.

- (1) A (2) B (3) C (4) D

58. **Assertion :-** For an isolated system, internal energy remains constant. [AIIMS 2018]

Reason :- Heat is exchanged through walls of isolated system.

- (1) A (2) B (3) C (4) D

59. **Assertion :-** Between states (P_1, V_1, T_1) and (P_2, V_2, T_2) , work done in two step process is more than work done in one step process. [AIIMS 2018]

Reason :- Work done is a path function.

- (1) A (2) B (3) C (4) D

60. **Assertion :-** In process of throttling, temperature neither increases nor decreases. [AIIMS 2018]

Reason :- Throttling is an isothermal process.

- (1) A (2) B (3) C (4) D

61. **Assertion :-** In an adiabatic process, work done by an ideal gas equals change in its internal energy.

Reason :- In an adiabatic process, temperature of system and surrounding remains same.

[AIIMS 2018]

- (1) A (2) B (3) C (4) D

62. **Assertion :-** On increasing temperature, internal energy of monoatomic gas increases.

Reason :- Thermal energy increases only random motion of gas molecules. [AIIMS 2018]

- (1) A (2) B (3) C (4) D

63. **Assertion :-** In a thermoflask, inner surface is shiny.

Reason :- Shiny surface is a poor radiator of heat.

- (1) A (2) B (3) C (4) D

64. **Assertion :-** In free expansion of an ideal gas internal energy is constant. [AIIMS 2018]

Reason :- No heat is added and no work is done in free expansion.

- (1) A (2) B (3) C (4) D

65. Assertion :- Pressure in a real gas is less than that in an ideal gas. [AIIMS 2018]

Reason :- Force of attraction is non-zero in a real gas.

- (1) A (2) B (3) C (4) D

66. Assertion :- Food is cooked with difficulty at mountain peak. [AIIMS 2018]

Reason :- Water boils at high temperature there.

- (1) A (2) B (3) C (4) D

67. Assertion :- COP of a refrigerator can't be greater than unity. [AIIMS 2018]

Reason :- Some work is done by working substance.

- (1) A (2) B (3) C (4) D

68. Assertion :- In phase transition from liquid to vapour, potential energy increases. [AIIMS 2018]

Reason :- Temperature is constant during phase transition.

- (1) A (2) B (3) C (4) D

EXERCISE-IV (Assertion & Reason)

ANSWER KEY

Que.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15
Ans.	1	1	1	1	1	3	1	1	1	1	3	1	1	1	3
Que.	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
Ans.	1	1	1	3	3	2	1	1	1	2	1	4	4	1	2
Que.	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45
Ans.	1	1	1	1	4	3	2	2	1	2	1	1	1	4	1
Que.	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60
Ans.	4	1	1	3	1	1	4	3	1	3	4	3	3	4	3
Que.	61	62	63	64	65	66	67	68							
Ans.	3	3	1	1	1	3	4	2							