

ARJUNA (NEET)

Classification of Elements & Periodicity in Properties

DPP-05

1. The first ionisation potentials of Na, Mg, Al and Si are in the order
 (A) $\text{Na} < \text{Mg} > \text{Al} < \text{Si}$
 (B) $\text{Na} < \text{Mg} < \text{Al} > \text{Si}$
 (C) $\text{Na} > \text{Mg} > \text{Al} > \text{Si}$
 (D) $\text{Na} > \text{Mg} > \text{Al} < \text{Si}$
2. Which among the following elements has the highest value for third ionisation energy?
 (A) Mg (B) Al
 (C) Na (D) Ar
3. Which of the following configuration is associated with the biggest jump between first and second ionization energy?
 (A) $1s^2 2s^2 2p^5$ (B) $1s^2 2s^2 2p^6 3s^1$
 (C) $1s^2 2s^2 2p^4$ (D) $1s^2 2s^1$
4. A sudden large jump between the values of second and third ionisation energies of an element would be associated with the electronic configuration
 (A) $1s^2, 2s^2, 2p^6, 3s^1$
 (B) $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$
 (C) $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$
 (D) $1s^2, 2s^2, 2p^6, 3s^2$
5. The element which has highest 2nd ionisation energy is
 (A) Na (B) Mg
 (C) Ca (D) Ar
6. In which of the following the energy change corresponds to first ionisation potential only:-
 (A) $\text{X}_{(\text{g})} \longrightarrow \text{X}_{(\text{g})}^{+} + \text{e}^{-}$
 (B) $\text{X}_{2(\text{g})} \longrightarrow \text{X}_{(\text{g})}^{+} + \text{e}^{-}$
 (C) $\text{X}_{(\text{s})} \longrightarrow \text{X}_{(\text{g})}^{+} + \text{e}^{-}$
 (D) $\text{X}_{(\text{aq})} \longrightarrow \text{X}_{(\text{aq})}^{+} + \text{e}^{-}$
7. The correct order of decreasing second ionization energy of Li, Be, Ne, C, B
 (A) $\text{Ne} > \text{B} > \text{Li} > \text{C} > \text{Be}$
 (B) $\text{Li} > \text{Ne} > \text{C} > \text{B} > \text{Be}$
 (C) $\text{Ne} > \text{C} > \text{B} > \text{Be} > \text{Li}$
 (D) $\text{Li} > \text{Ne} > \text{B} > \text{C} > \text{Be}$
8. In which of the following element has highest value of ionisation energy
 (A) Ti (B) Zr
 (C) Hf (D) None of these
9. Minimum first ionisation energy is shown by which electronic configuration:-
 (A) $1s^2 2s^2 2p^5$
 (B) $1s^2 2s^2 2p^6 3s^2 3p^2$
 (C) $1s^2 2s^2 2p^6 3s^1$
 (D) $1s^2 2s^2 2p^6$
10. The energy needed to remove one electron from unipositive ion is abbreviated as:-
 (A) 1st I.P. (B) 3rd I.P.
 (C) 2nd I.P. (D) 1st E.A.

ANSWER KEY

1. (A)
2. (A)
3. (D)
4. (D)
5. (A)
6. (A)
7. (D)
8. (C)
9. (C)
10. (C)



Note - If you have any query/issue



Mail us at support@physicswallah.org
