

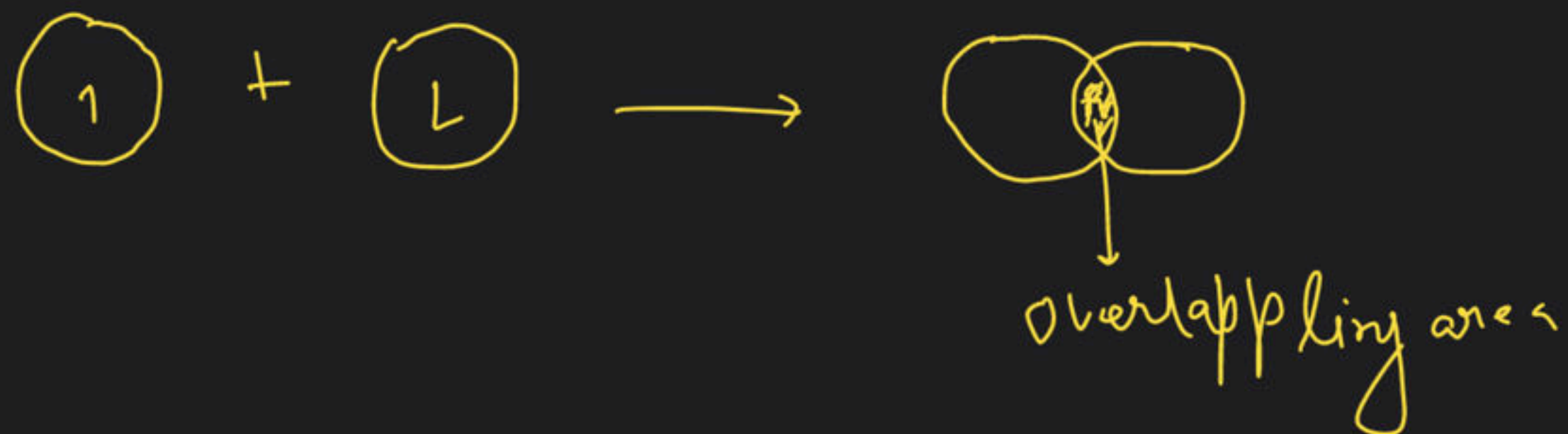


Bond Strength, Catenation Atomicity

Course on Chemical Bonding for Class XI 2023

V.B.T \rightarrow (Valence bond theory)

\rightarrow Covalent bond formation through overlapping



overlapping area \uparrow Covalent bond strength \uparrow

Ques Which of the following overlapping is correct

① $(12) + (12) \rightarrow \times$

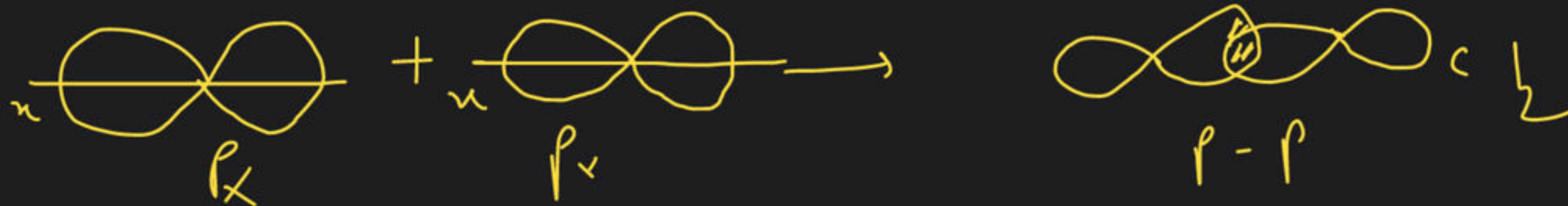
② $(12) + (1) \rightarrow \times$

③ $() + () \rightarrow \times$

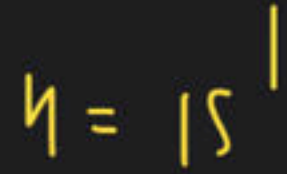
~~④~~ none

type of overlapping $\begin{cases} \text{axial [Head to Head]} \\ \text{Side Ways [Co Metral]} \end{cases}$

① axial overlapping [σ bond]
intermolecular axis σ

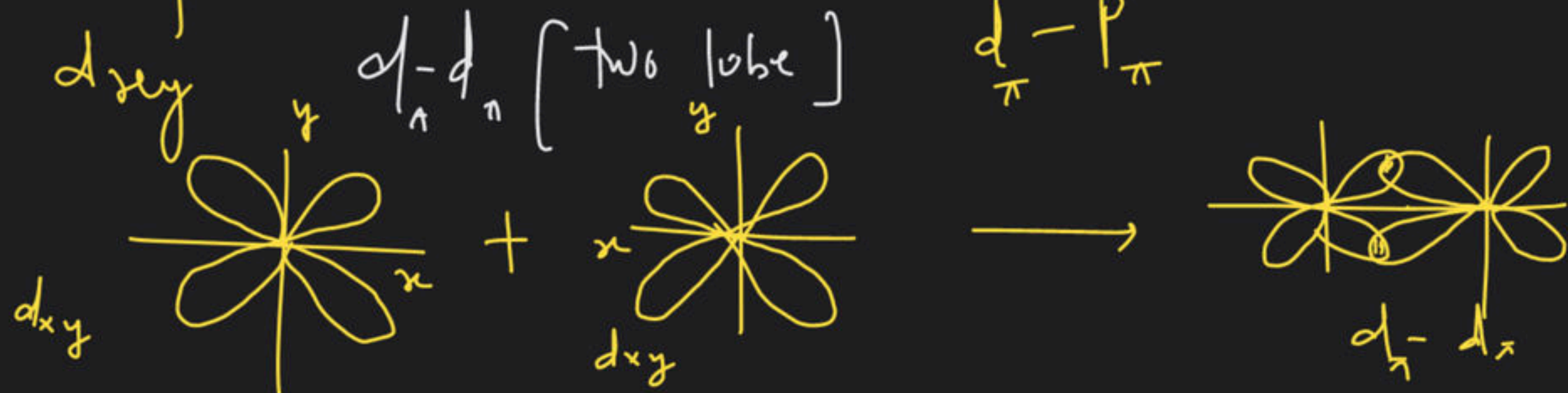
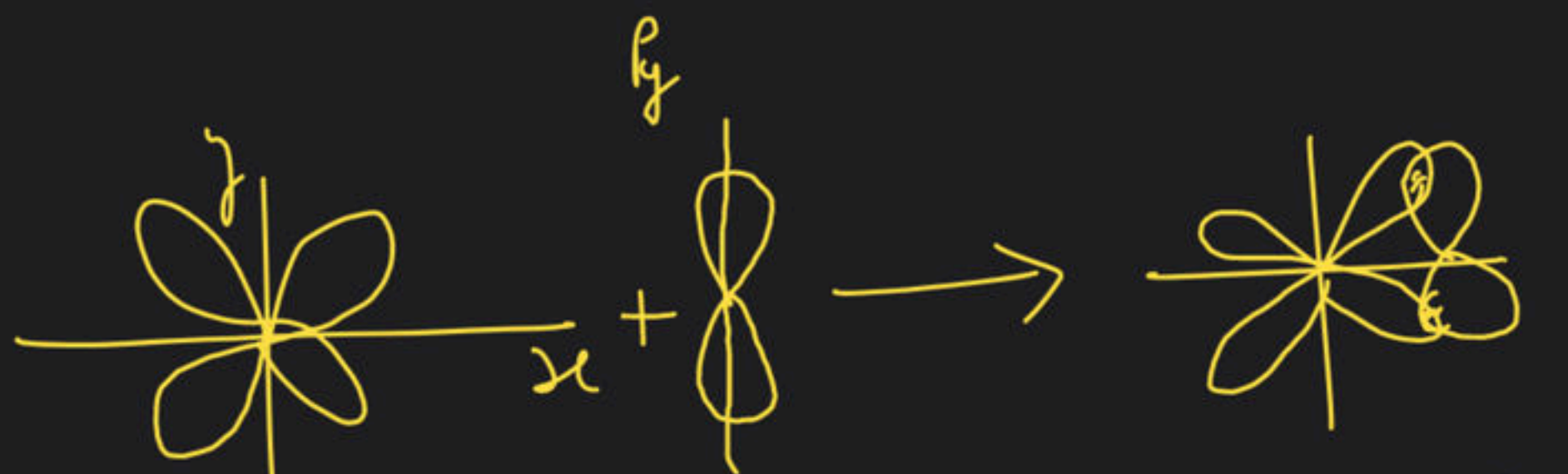


one which of the following overlapping is s-s

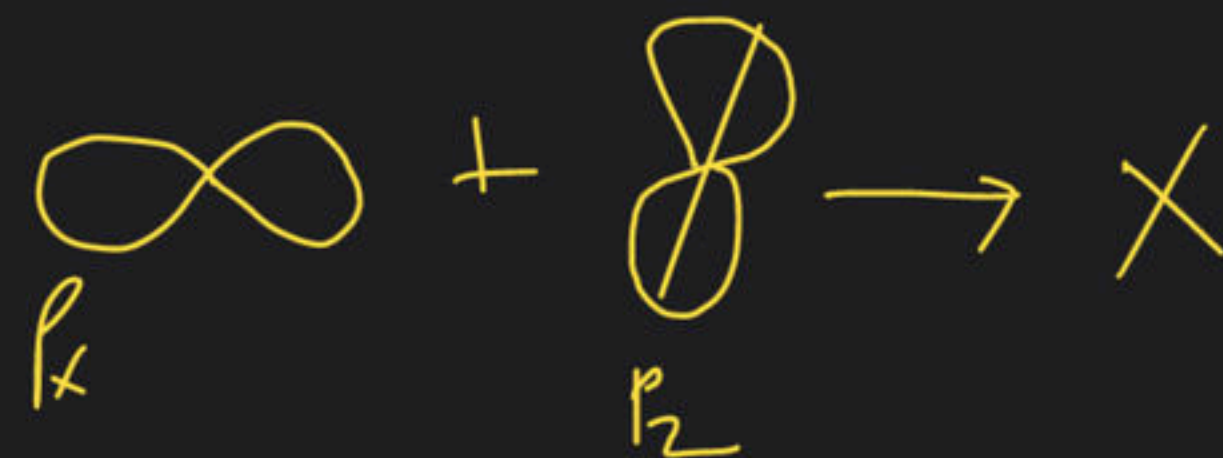
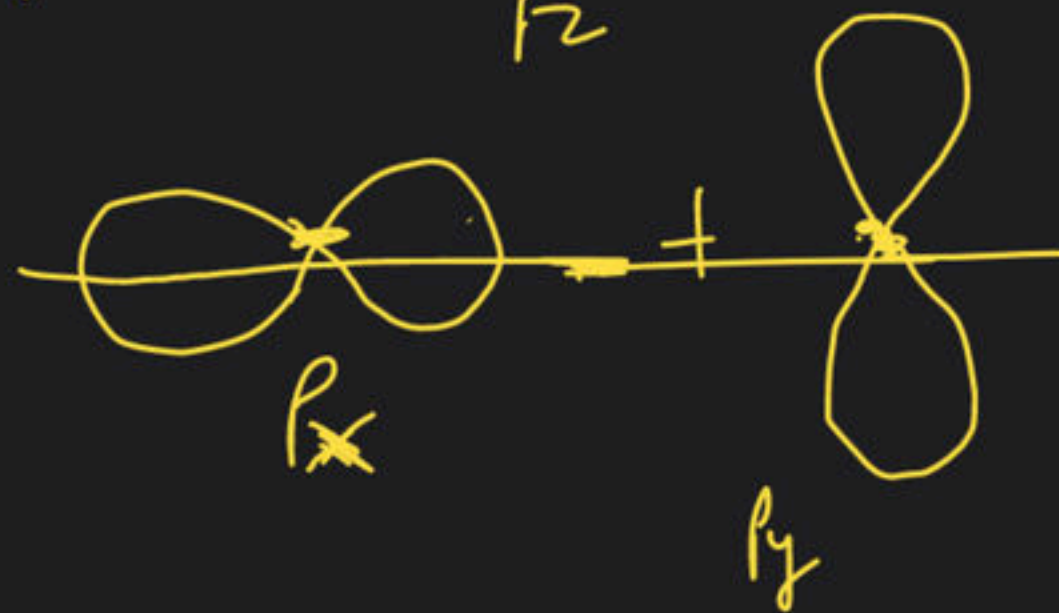
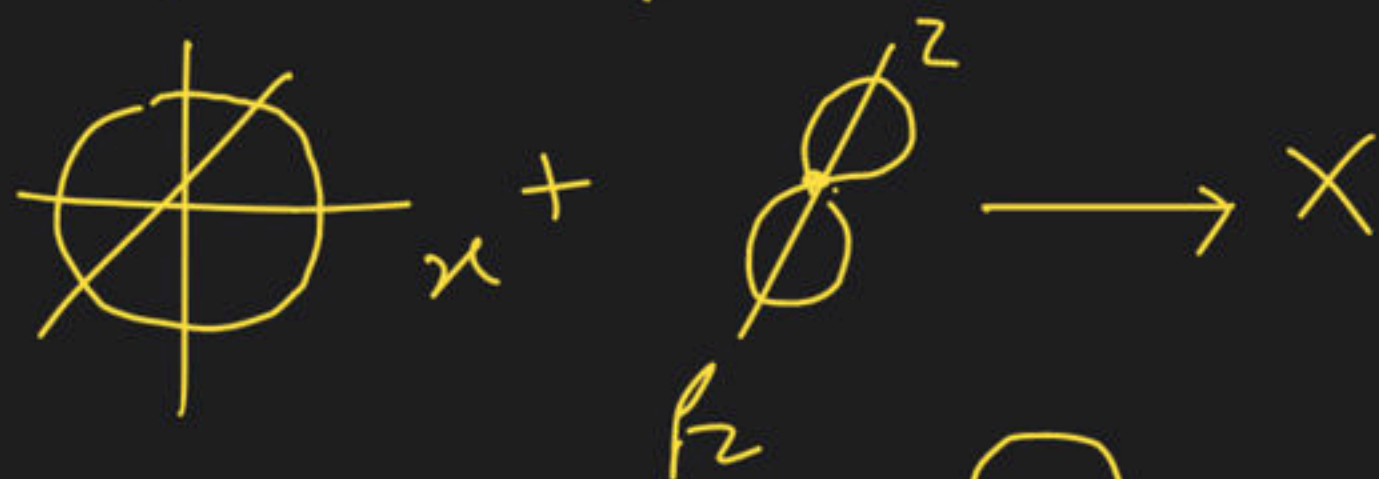
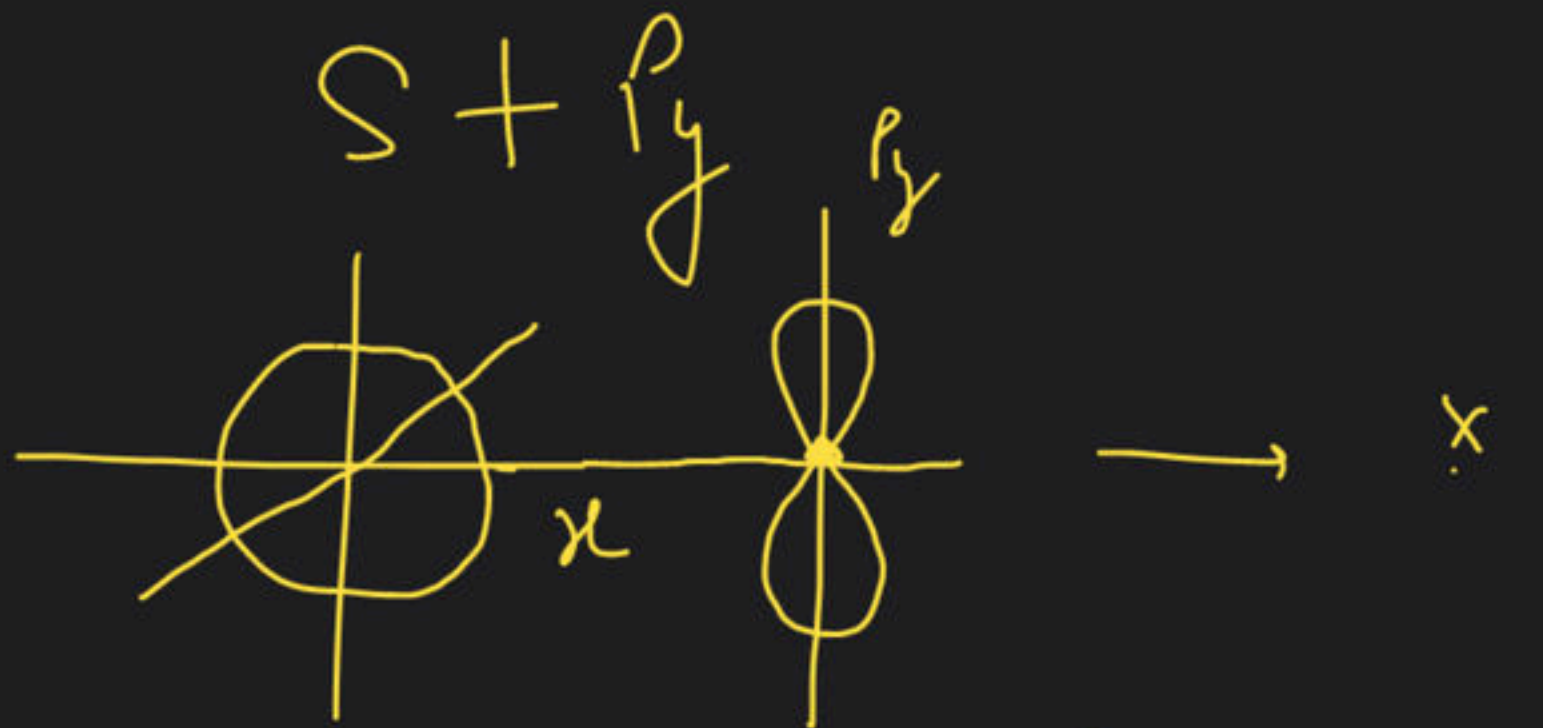


Sideways (π bond)

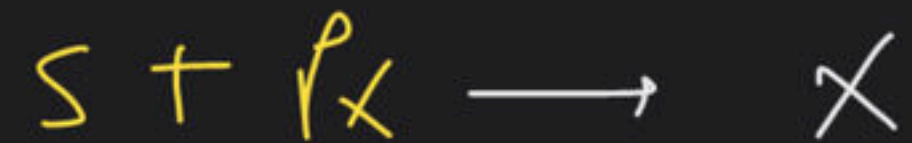
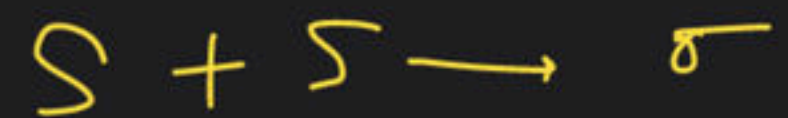
internuclear axis x



If x is internuclear axis



if z is internuclear axis



$$x \quad p_x \perp p_x = \tau$$

$$p_y \perp p_y = \pi$$

$$p_1 \perp p_2 = \pi$$

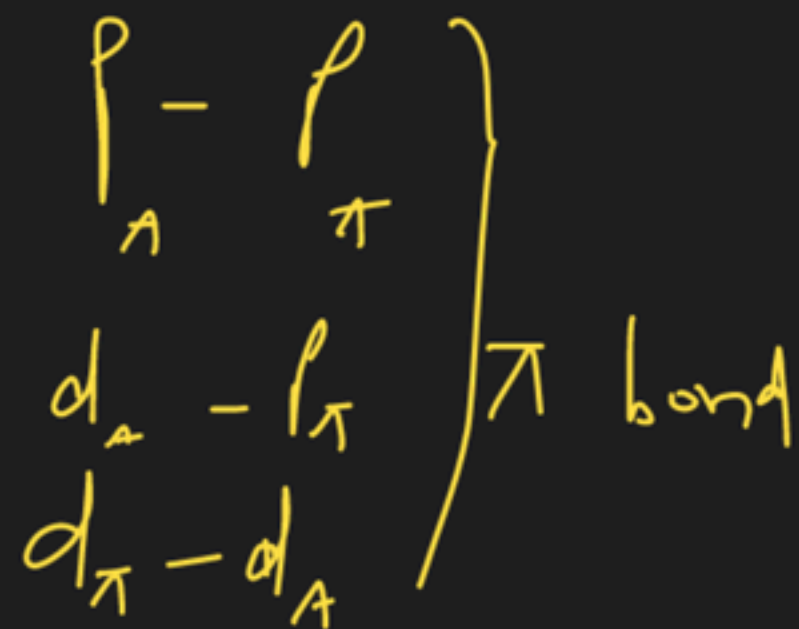
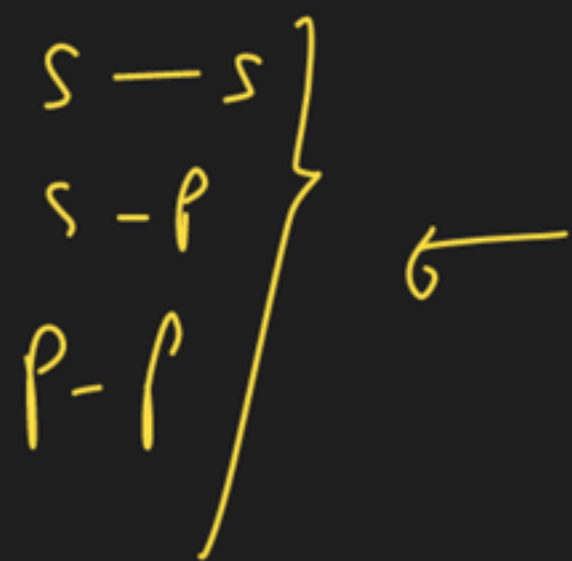
$$p_x \perp p_y$$

$$p_1 \perp p_2$$

$$p_1 \perp p_x$$

$$0 + \infty \rightarrow \checkmark$$

$$\left. \begin{array}{l} 0 + p_y \\ \perp p_1 \end{array} \right\} \checkmark$$



if y is interuclear axis:

$$S + S = \sigma$$

$$S + P_x = X$$

$$S + P_y = \sigma$$

$$S + P_z = X$$

$$P_x + P_y = X$$

$$P_x + P_z = X$$

$$P_x + P_x = \pi$$

$$P_y + P_y = \sigma$$

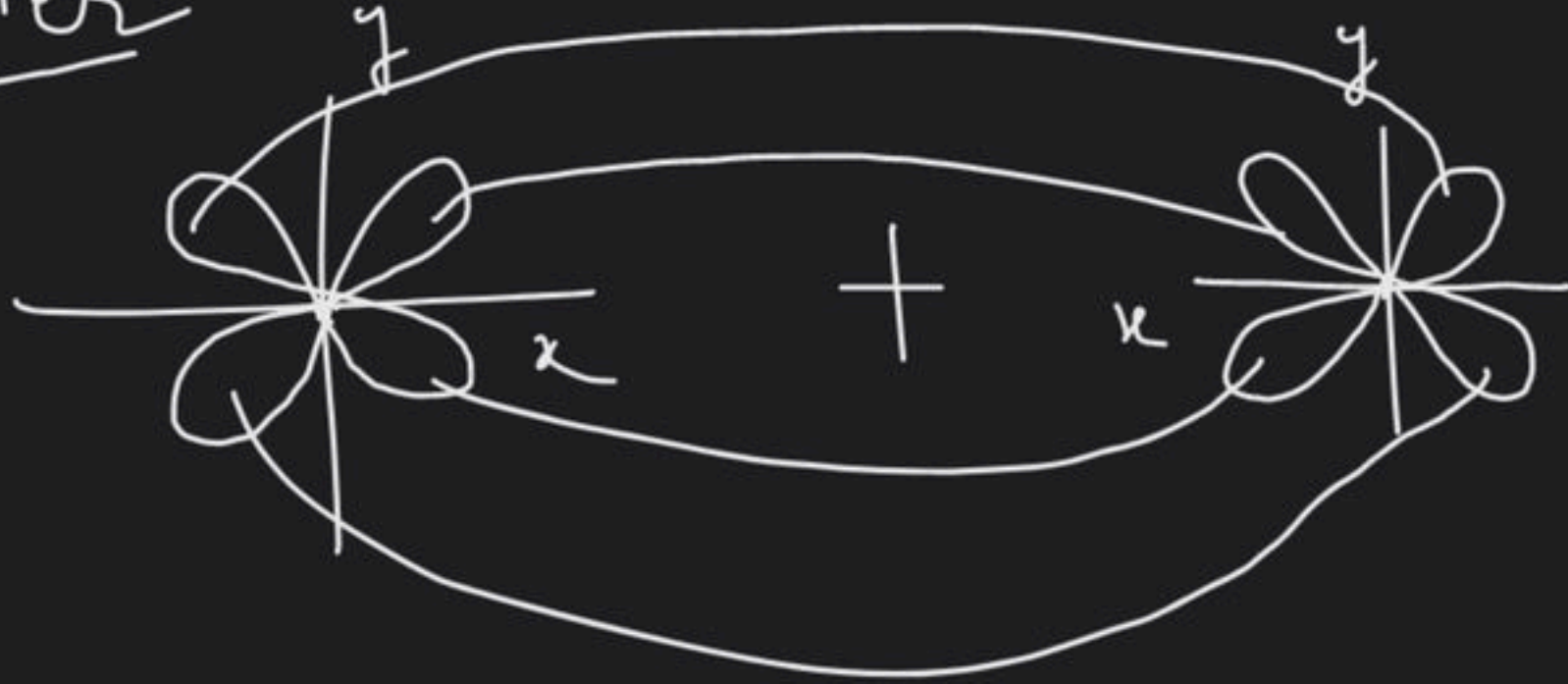
$$P_y + P_z = X$$

$$P_z + P_y = X$$

$$P_z + P_z = \pi$$

f-Bond \rightarrow [four lobe interaction]

if z is inter



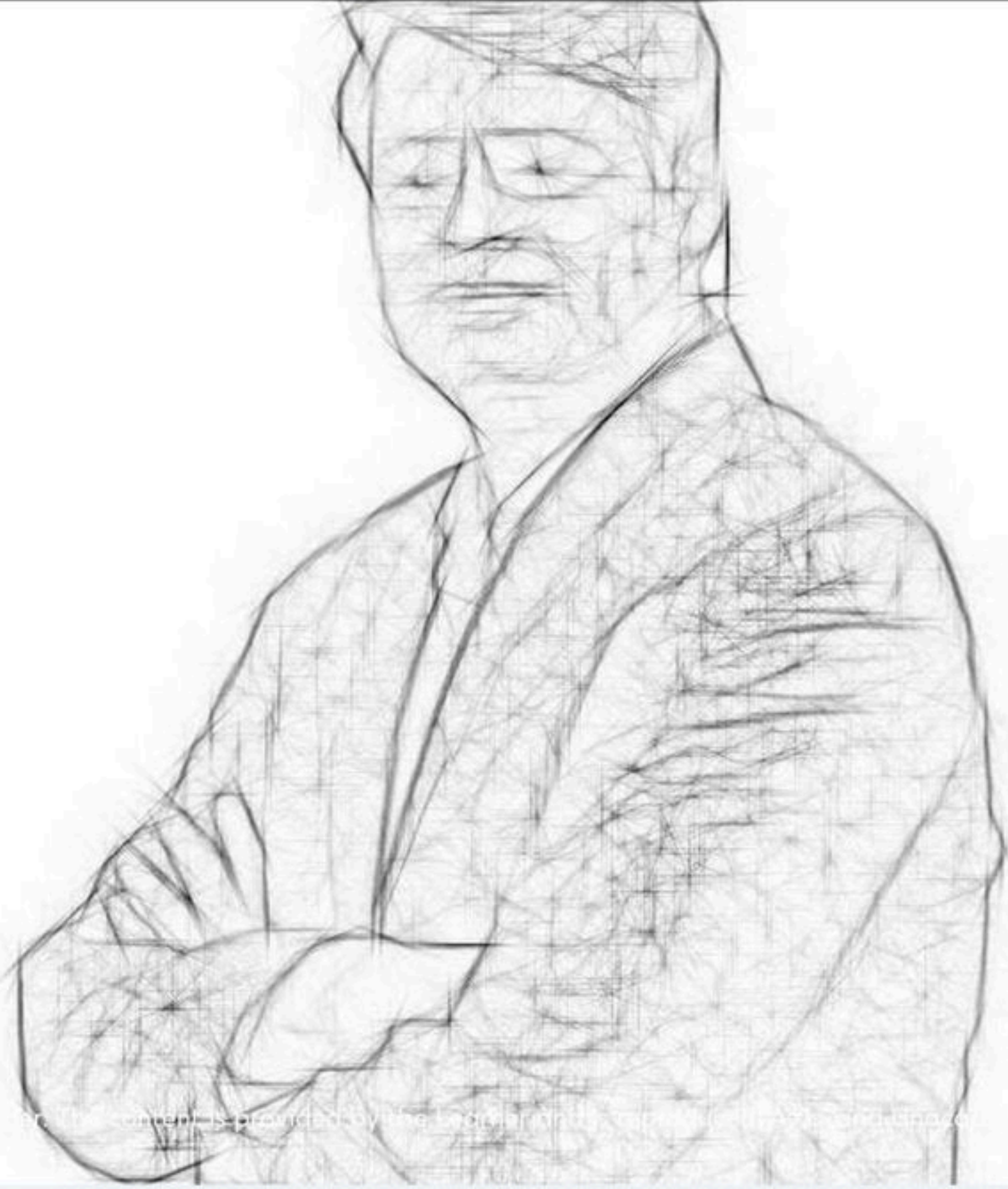
Note : all d-orbitals can form σ Bond but
not d_{z^2} because it has
two lobe interaction



Question

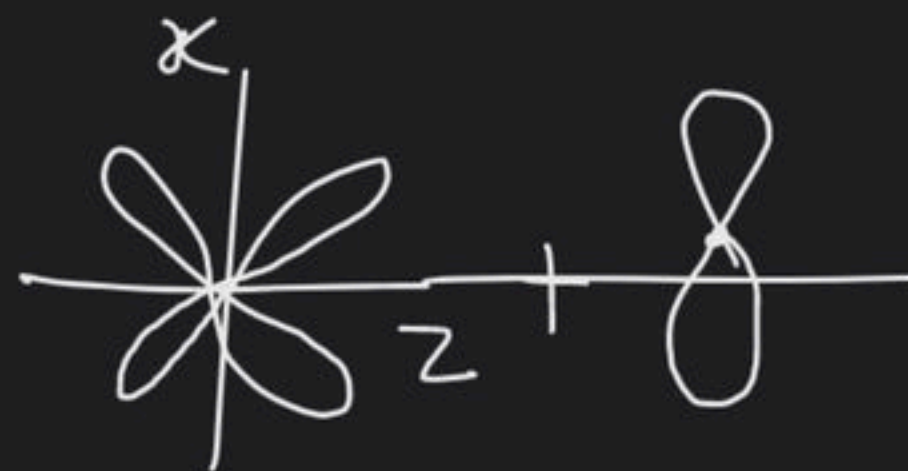
from Nikshey Ya...

sir aapke liye



and all liabilities with regards to the content

if z is internuclear axis



$$\cancel{d_{xy} + d_{xy} = \sigma}$$

$$\cancel{d_{xy} + d_{yz} = \pi}$$

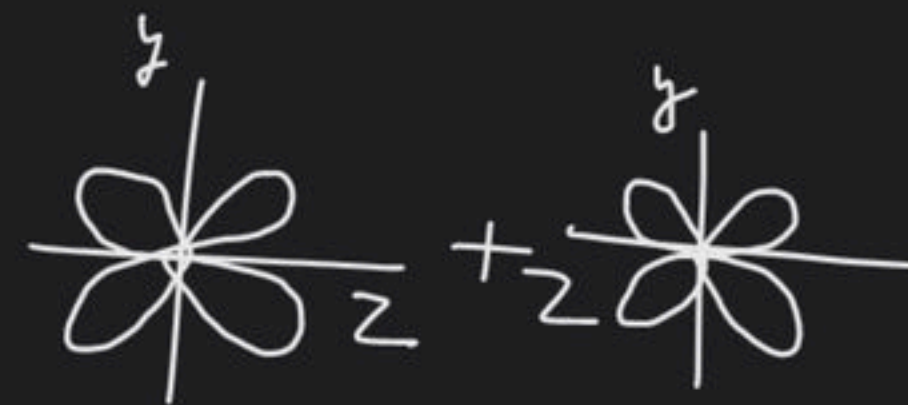
$$\cancel{d_{yz} + d_{zy} = \pi}$$

$$\cancel{d_{yz} + d_{zx} = \pi}$$

$$\cancel{s + d_{xy} = \sigma}$$

$$\cancel{d_{xz} + p_z = \sigma}$$

$$\cancel{d_{xz} + p_x = \pi}$$

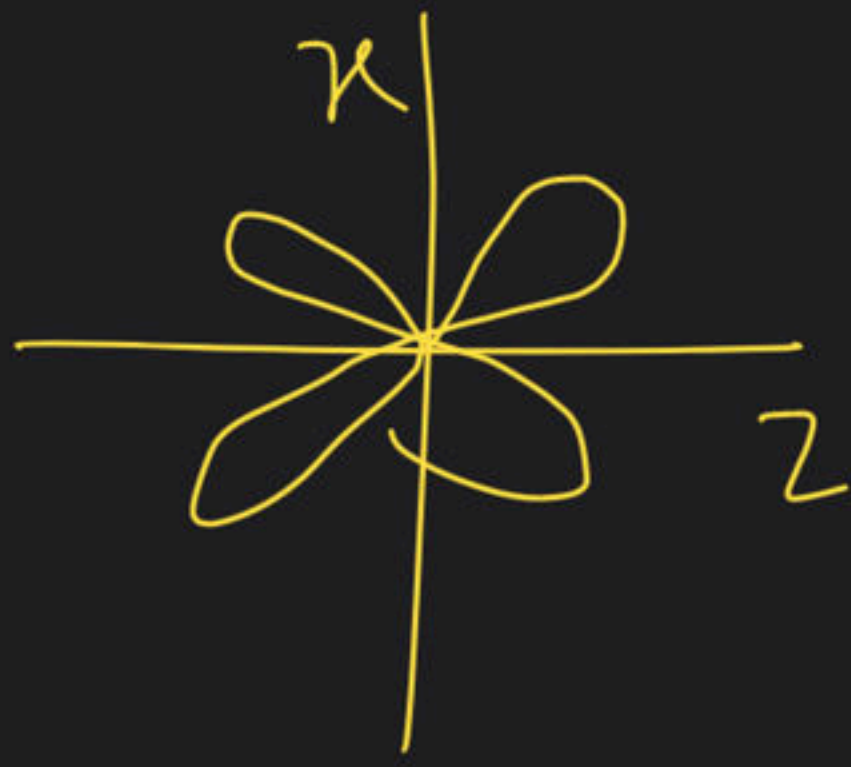


Grandpère

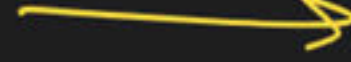
Ø bord

β's

1r - 1s 's



+

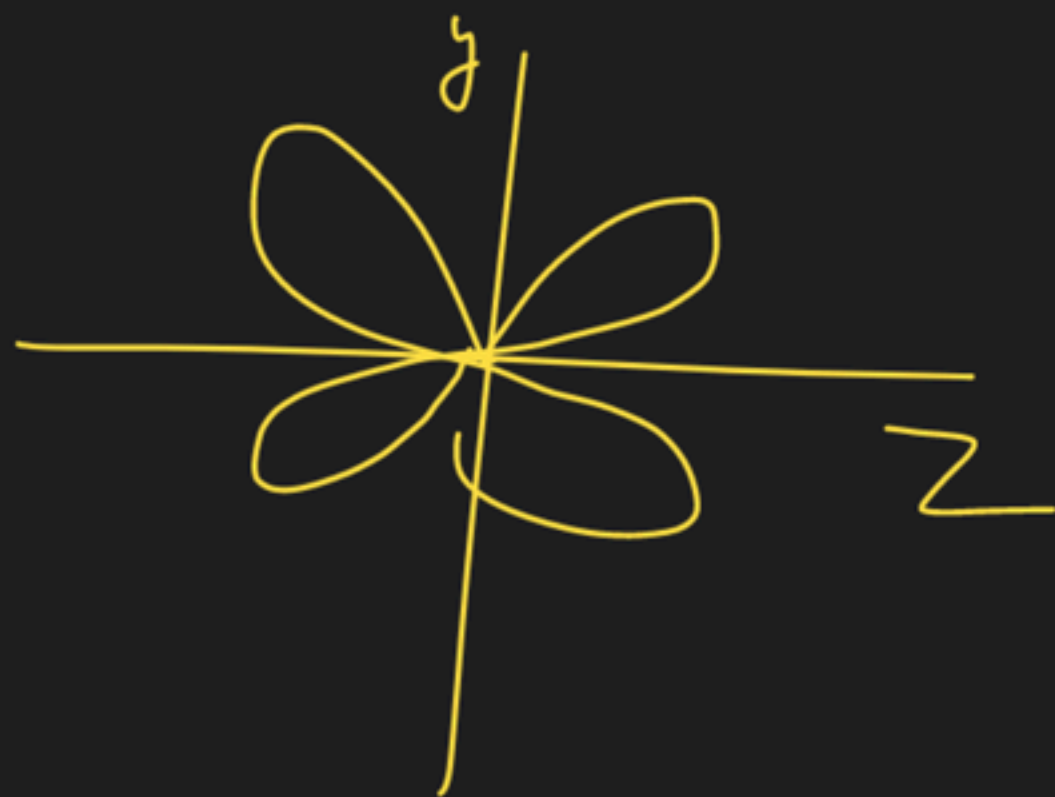


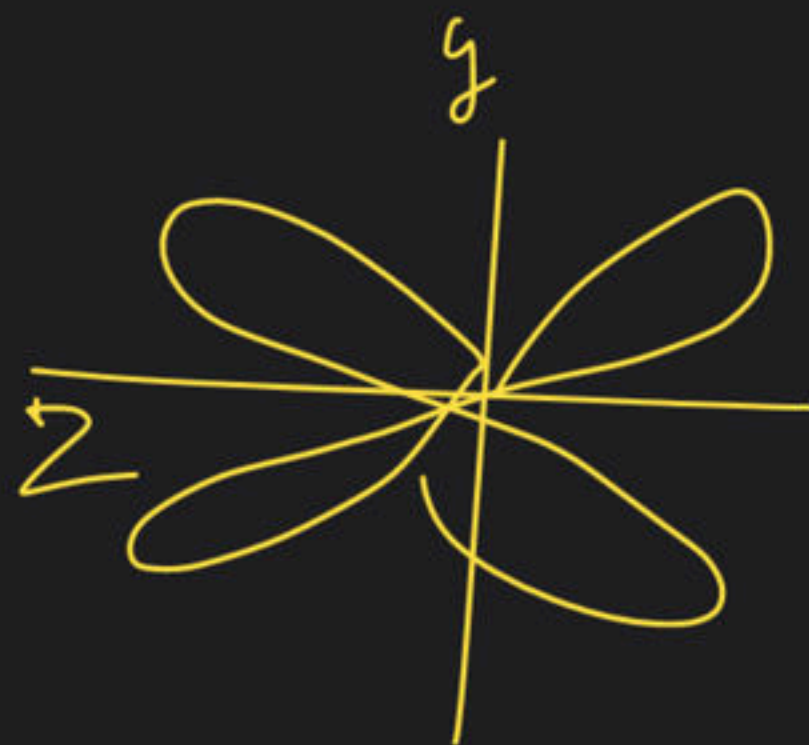
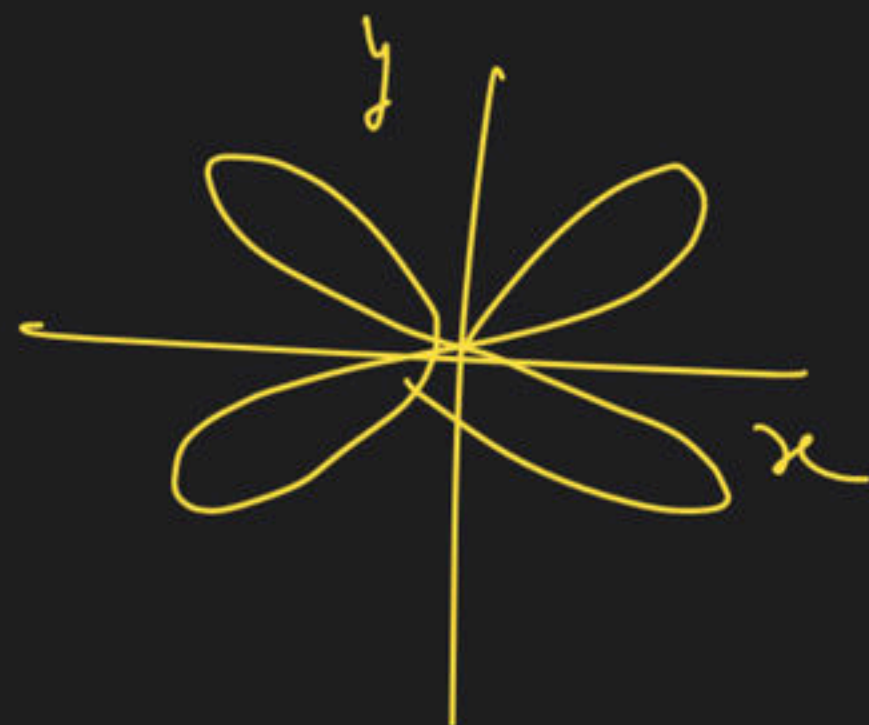


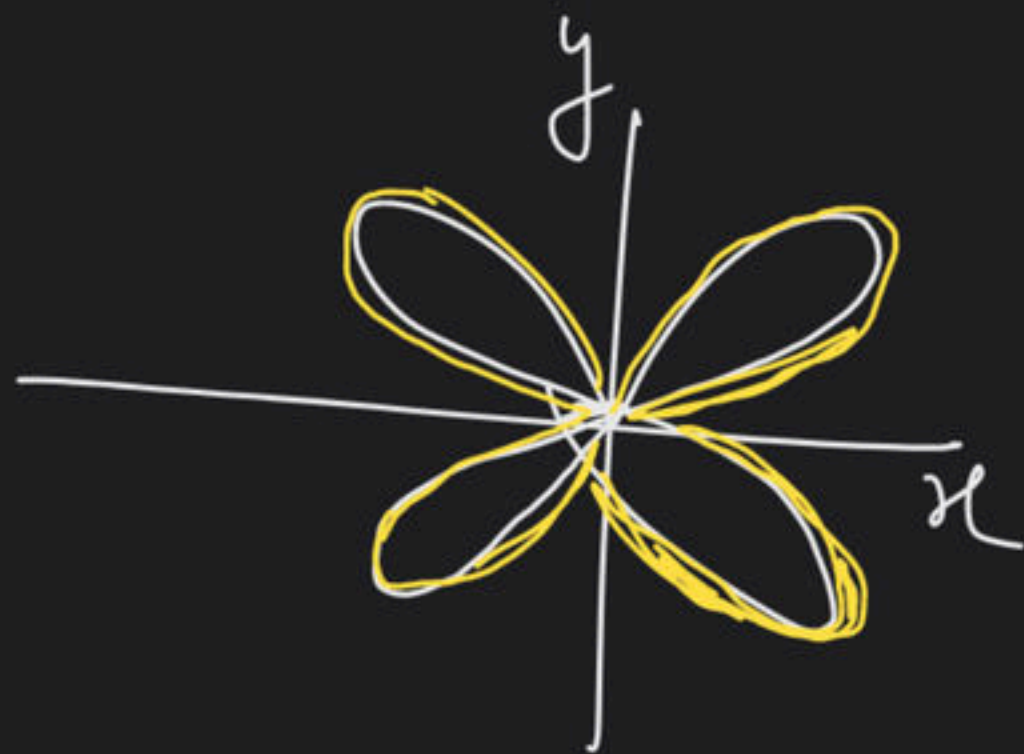


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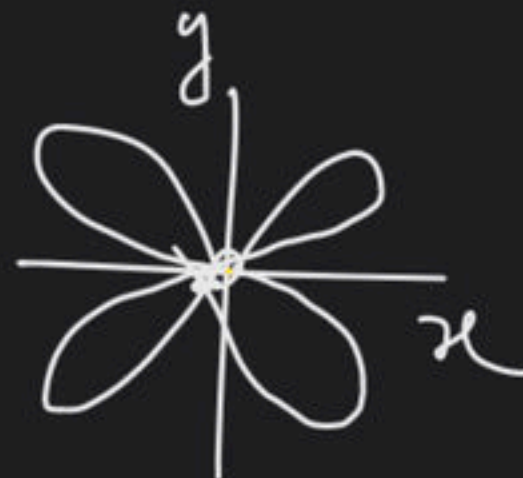








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Question

from Himanshu

unacademy


s-Bond \div [four lobe interaction]

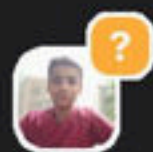
if z is inter



Note - all d-orbitals can form s bond but not d_{z^2} because it has two lobe interaction

LIVE



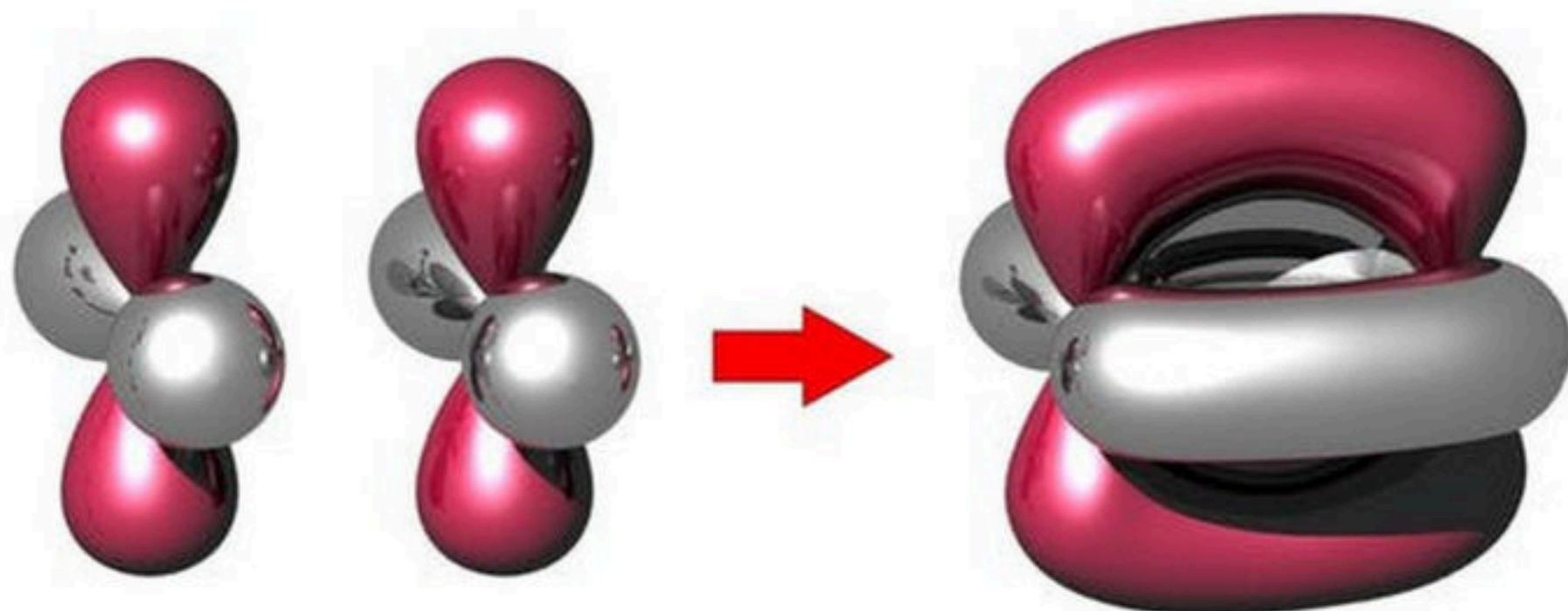


Question

from ARNAVGUPTA

delta bond

Delta Bond Type



$d + d$

$\delta d - \delta d$

