



VSEPR - III

Course on Chemical Bonding for Class XI 2023



Question

from Roshan

Happy teacher day sir in advance



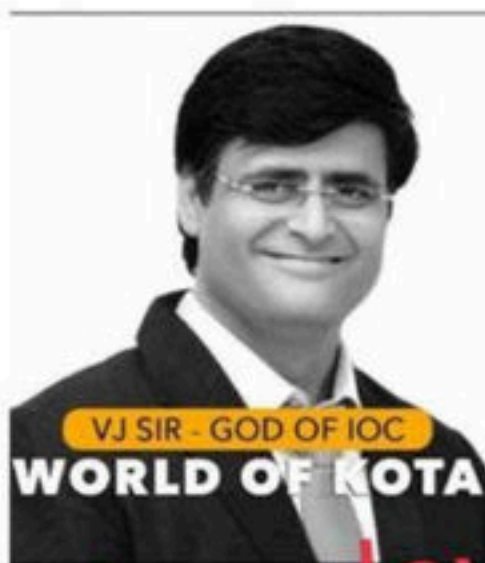
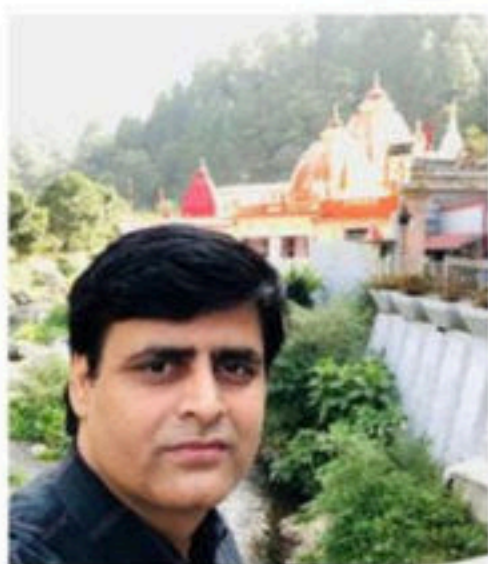
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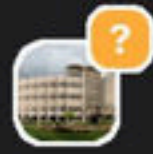
Question

from Adarsh

Please take me on vc



Happy teacher day
ENJOY YOUR LIFE
sir
Just breathing the air there was never a care.



Question

from Swarnim

Hapy Teachers day sir

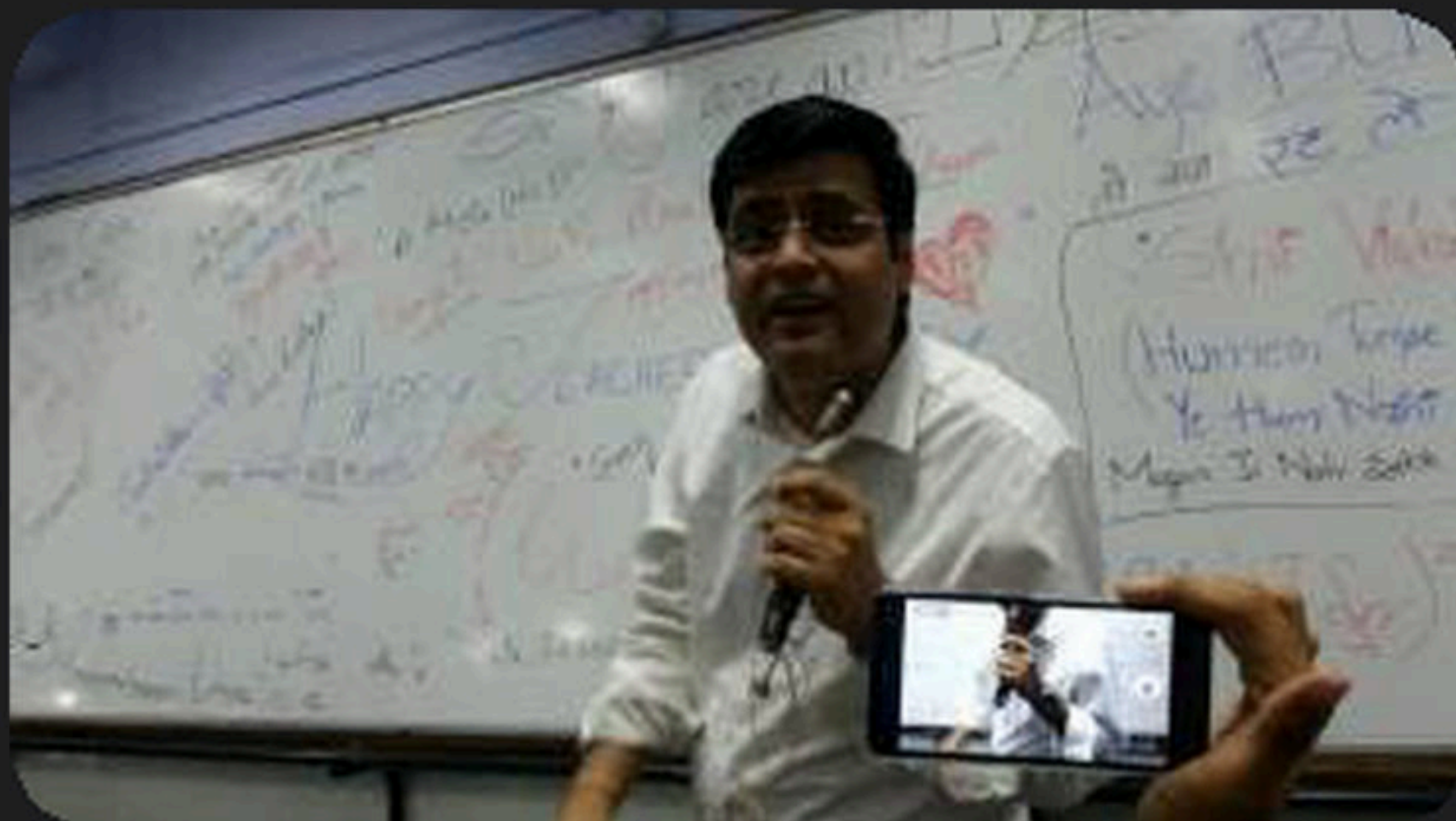


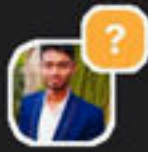


Question

from Kavya

Happy Teacher's day sir!





Question

from Akash

Happy teacher day sir





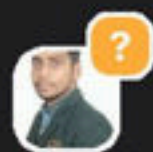
Question

from Ansh

Happy Teachers day in advance sirApna aashirwaad
banay rakhiyega....iss photo mein aap boht acche lag rhe







Question

from Ashutosh Y...

Pranaam sir happy teacher day advance

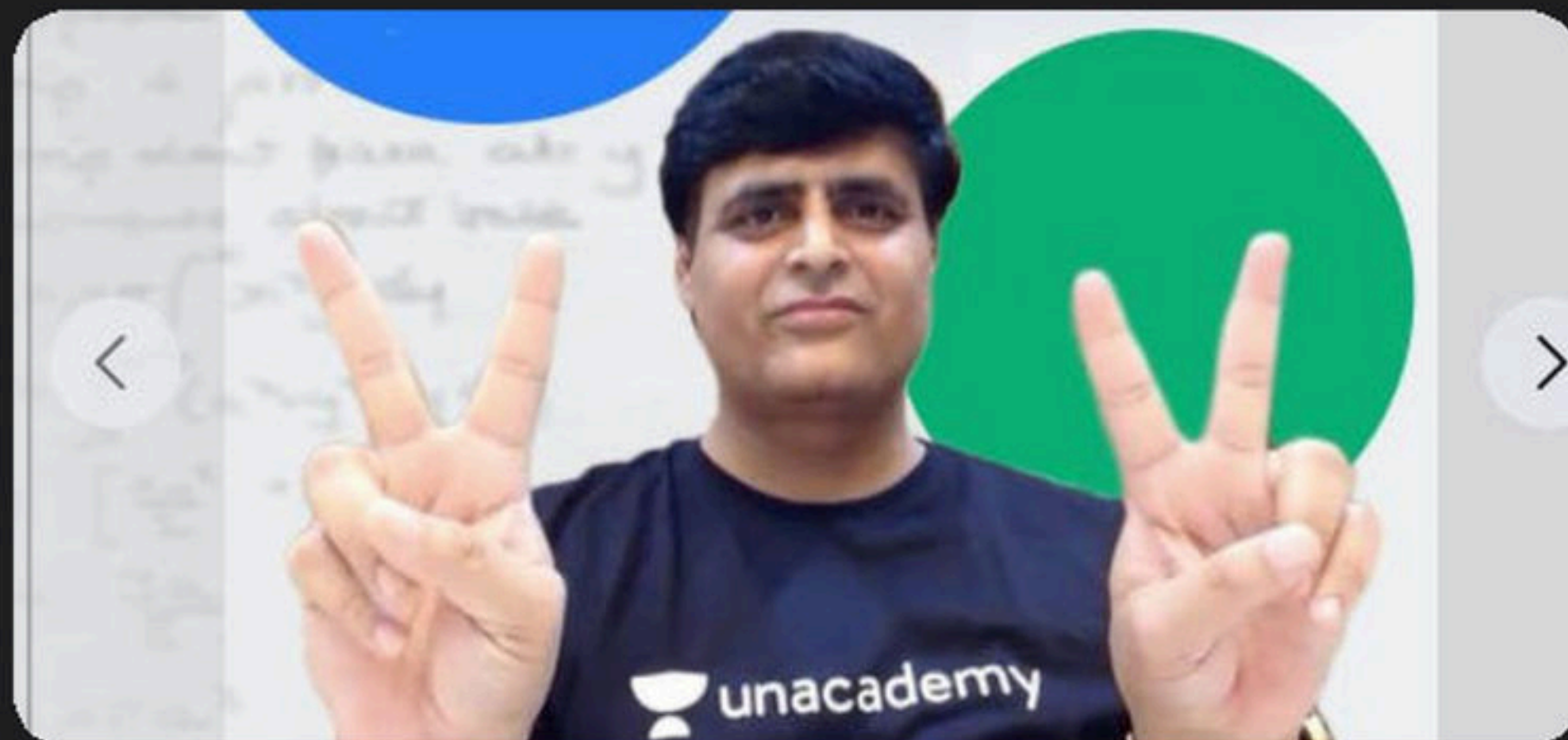




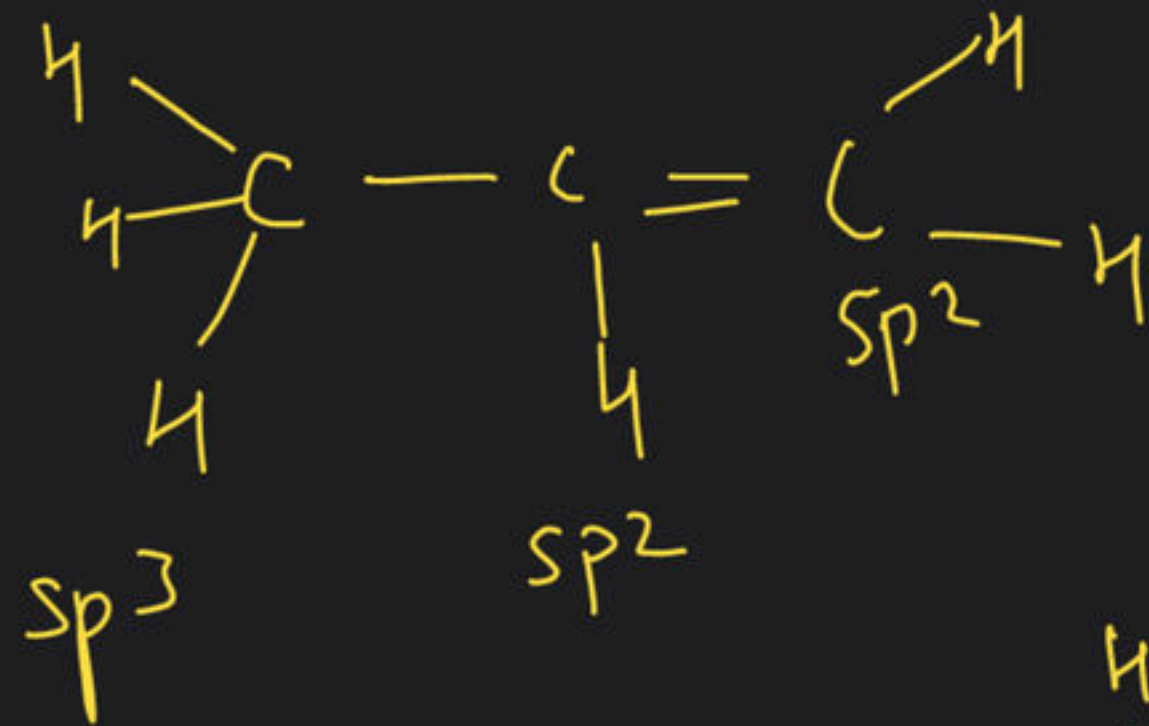
Question

from supreet

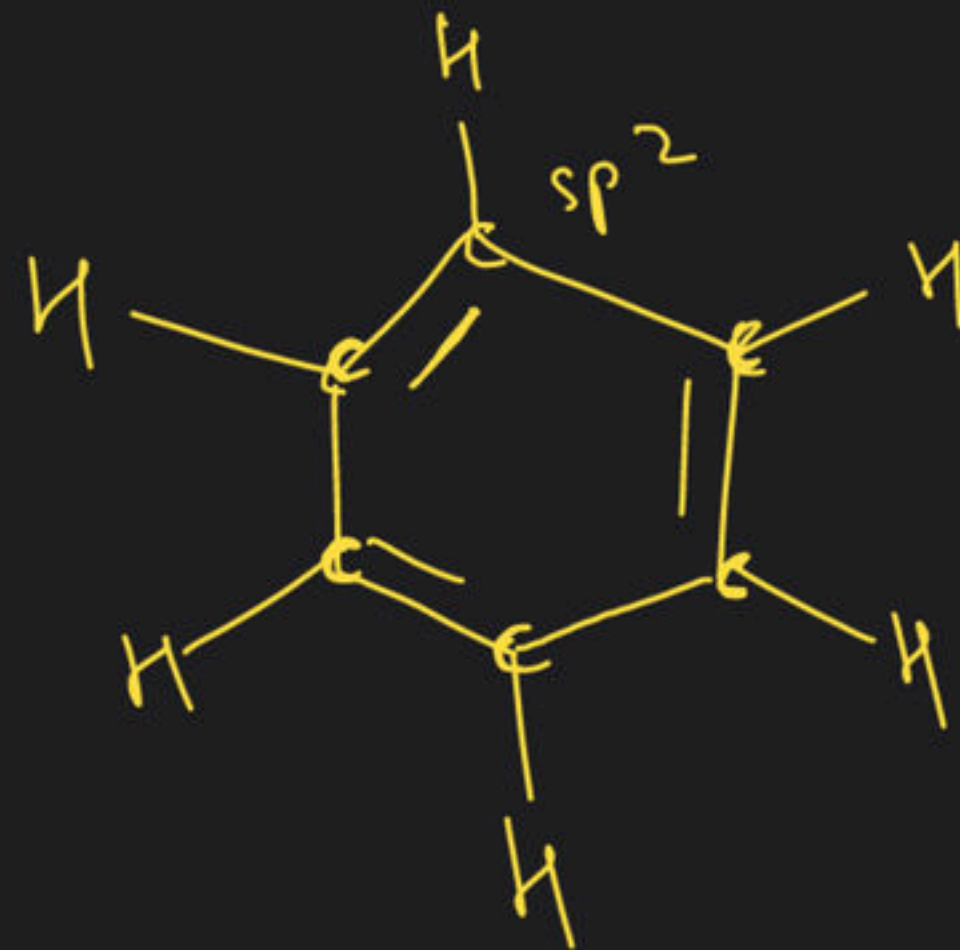
happy tr day



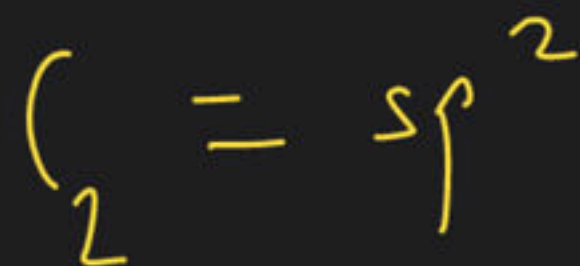
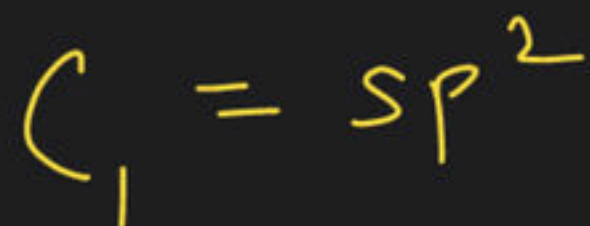
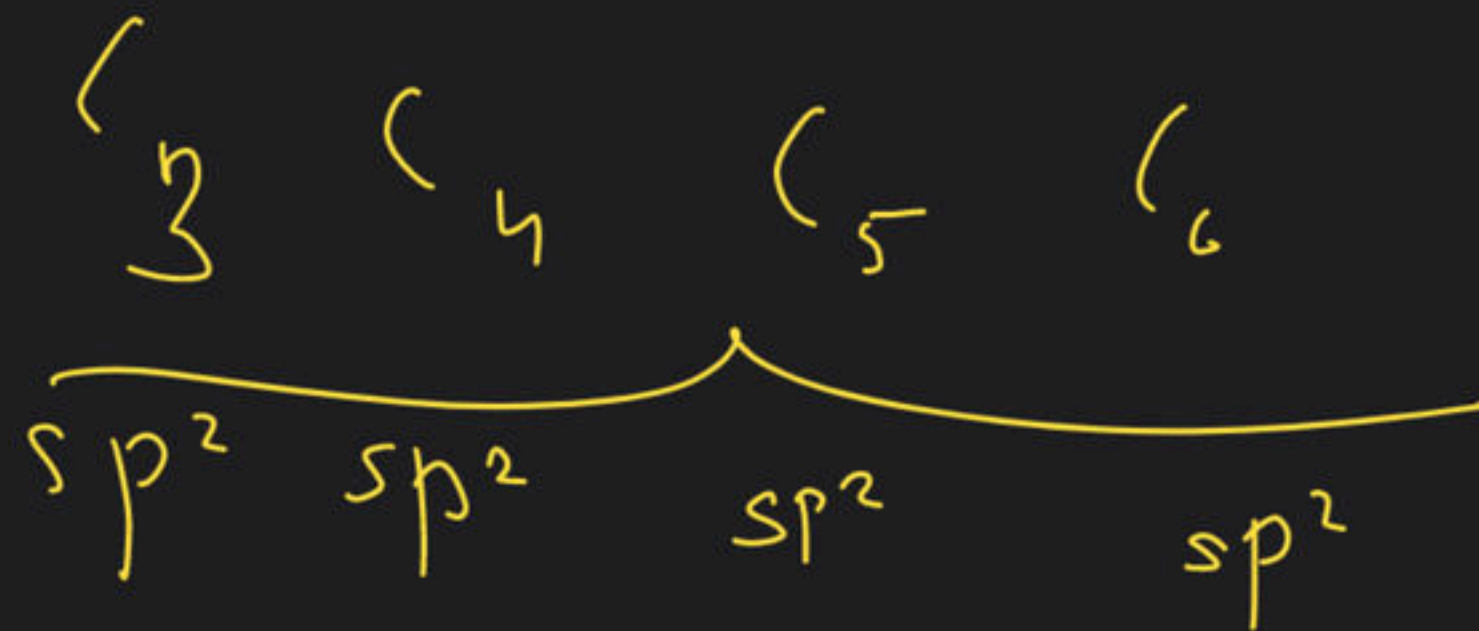
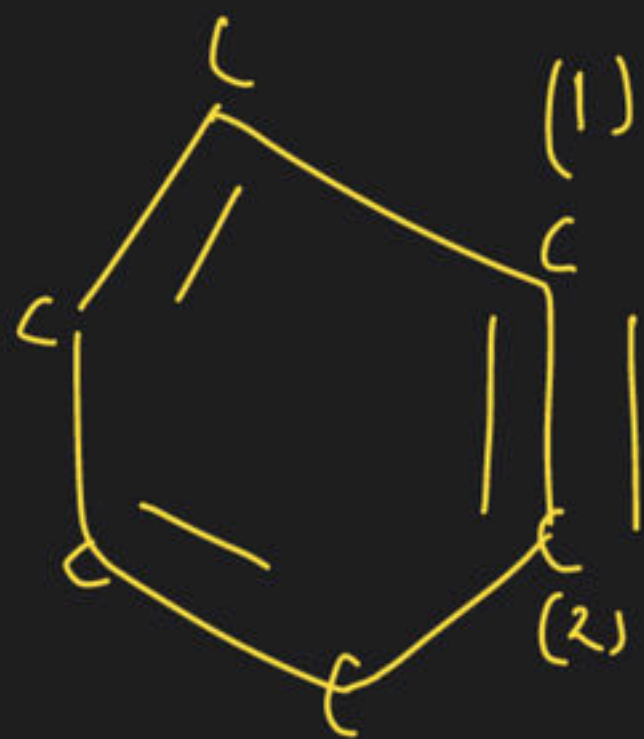
$\text{CH}_3 - \text{CH} = \text{CH}_2$ { all atoms are sp^2 hybridised in benzene. False



all Carbon atoms are sp^2 hyb. in benzene



benzene

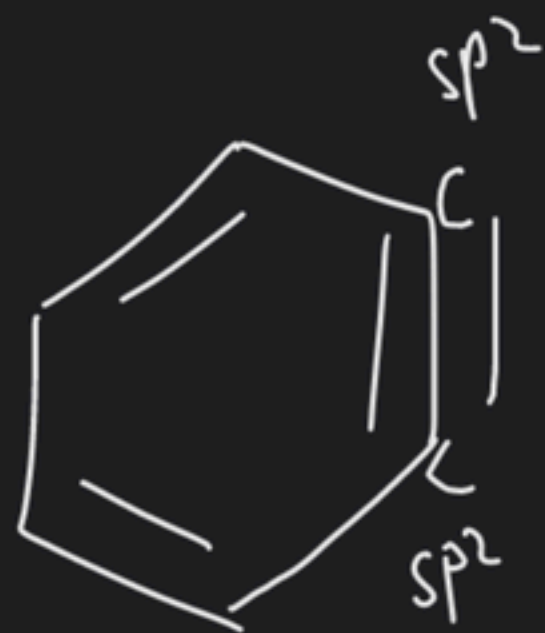


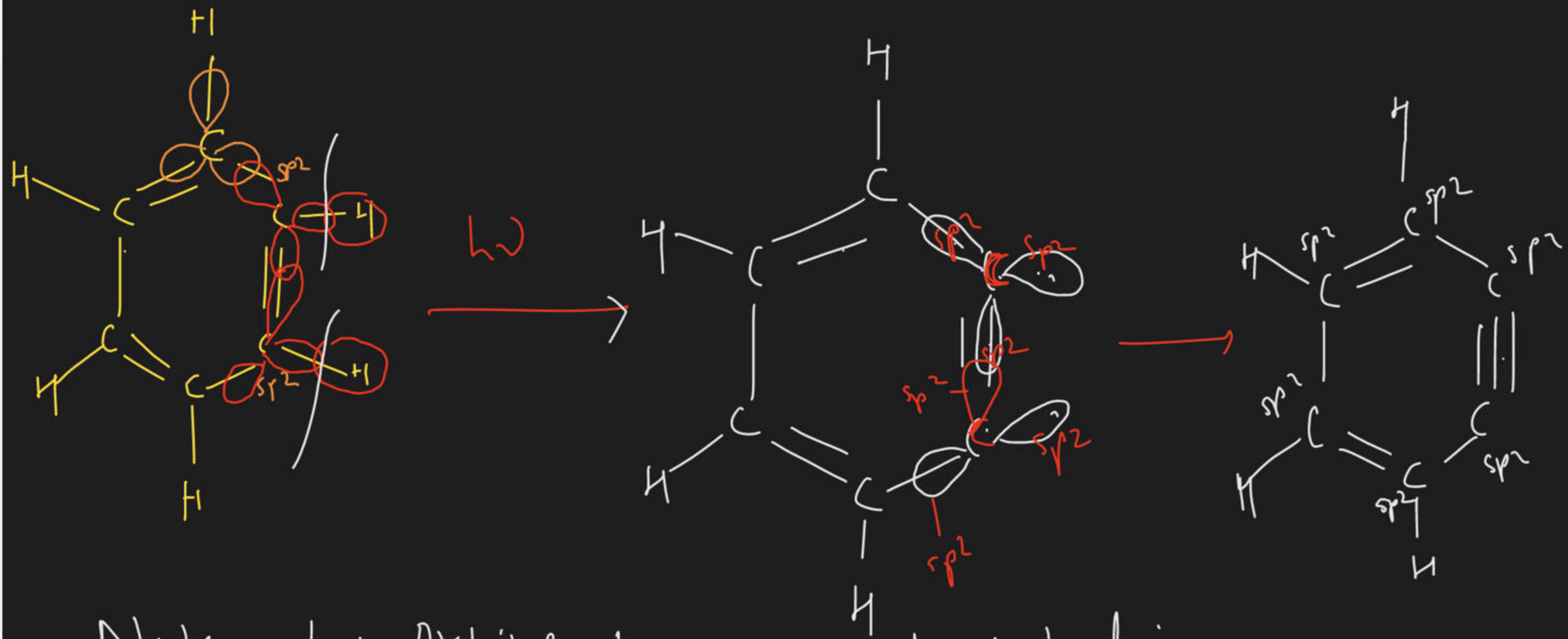
all
hyb. in benzene atoms have same
(F)

first hyb. followed by overlapping

$$\text{F.O.B.} = \text{no of } \sigma \text{ bonds} + \text{no of l.p.}$$

Trick

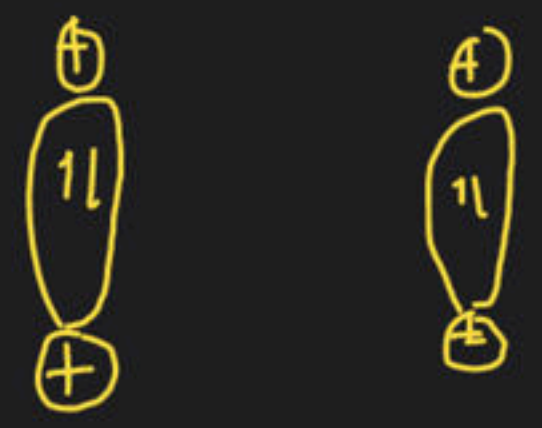
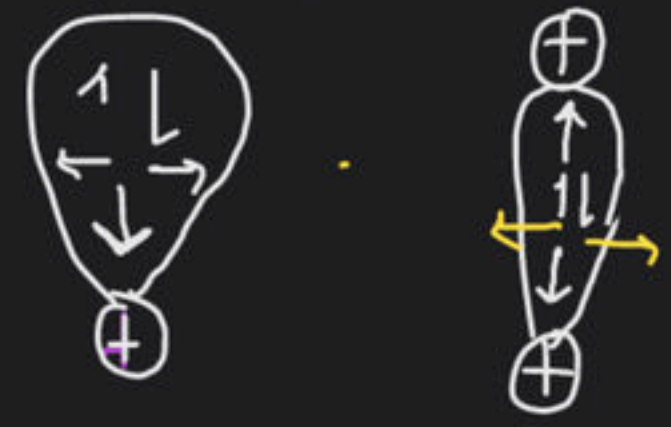
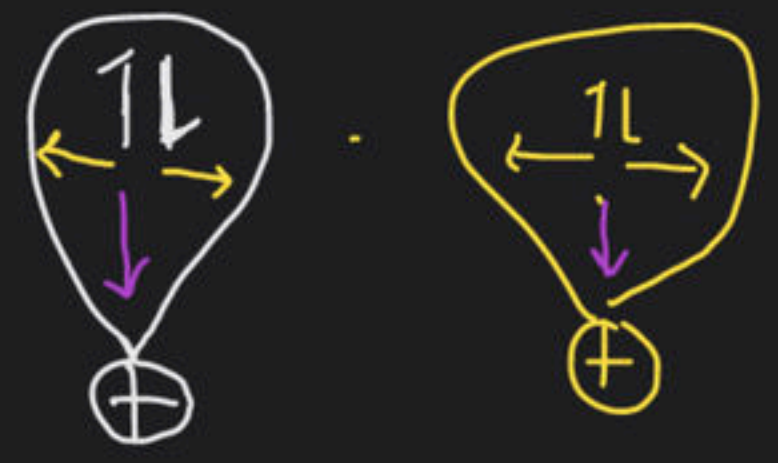




Note : sp^2 orbital forms σ bond but in benzene they form π bond also

V.S.E.P.R (minimum rep. theory) (rep. order)

$l.p - l.p > l.p - B.p > B.p - B.p$

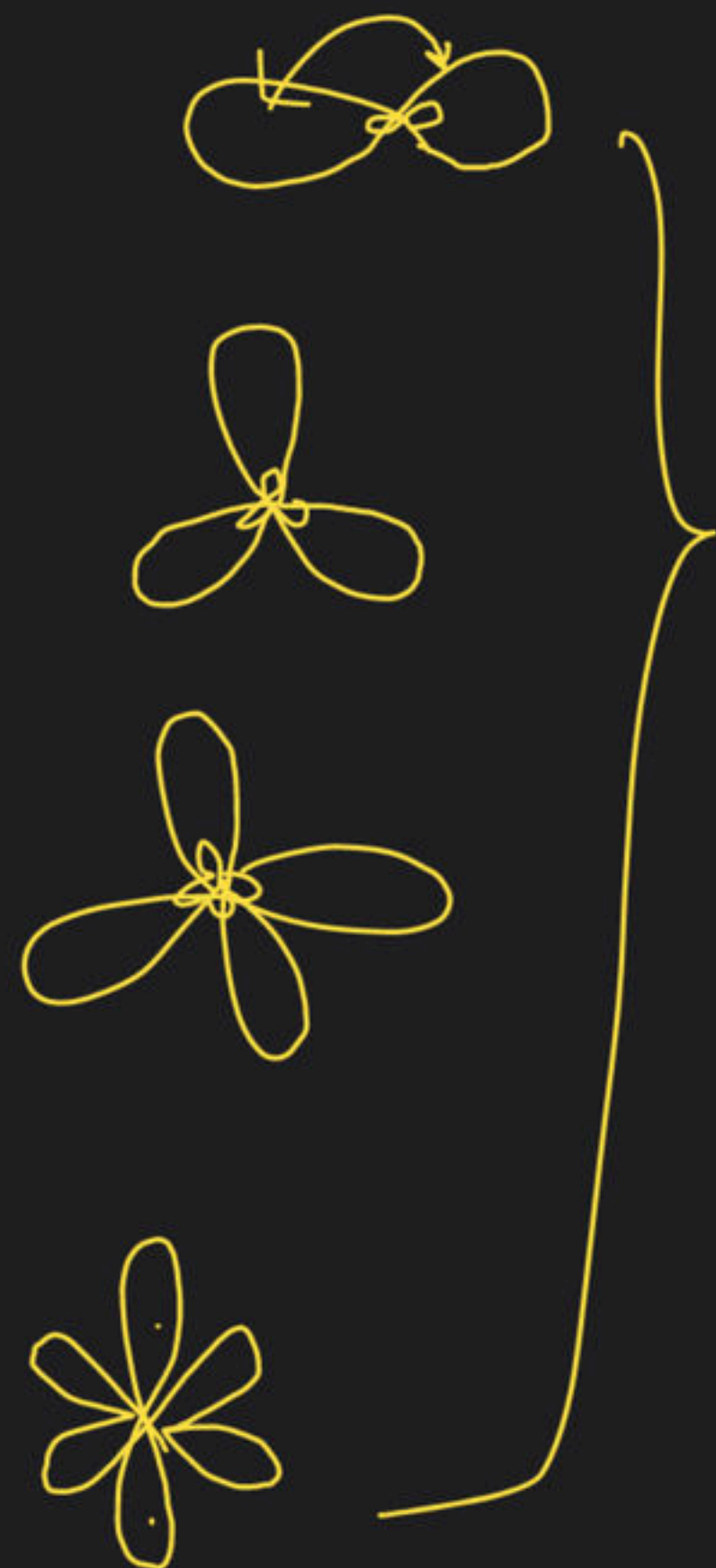


$A \rightarrow B$

all carbon atoms have same hyp. in benzyne



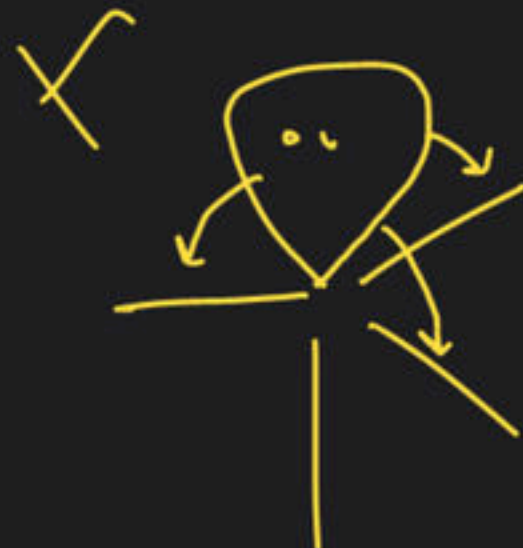
sp^2





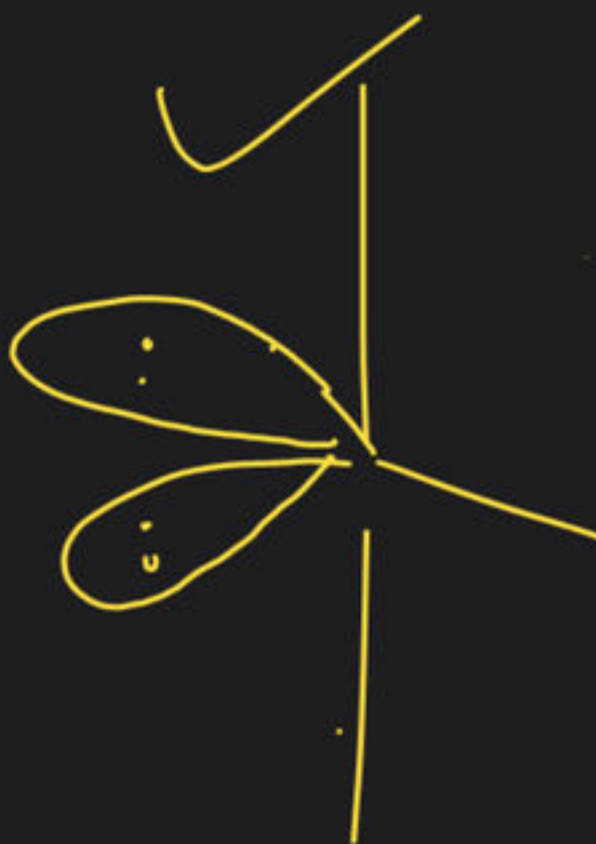
at go

$$l.p - B.p = 2$$



$$l.p - B.p = 3$$

See - Saw



at 50'

$$L.P - L.P = 0$$

$$L.P - B.P = 4$$

bent T shape



$$L.P - L.P = 1$$

$$L.P - B.P = 3$$



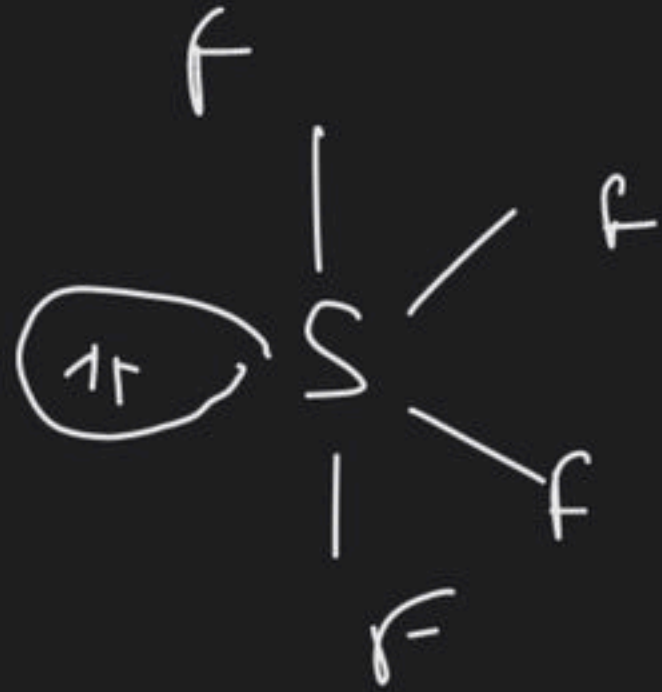
$$L.P - L.P = 0$$

$$L.P - B.P = \underline{6}$$

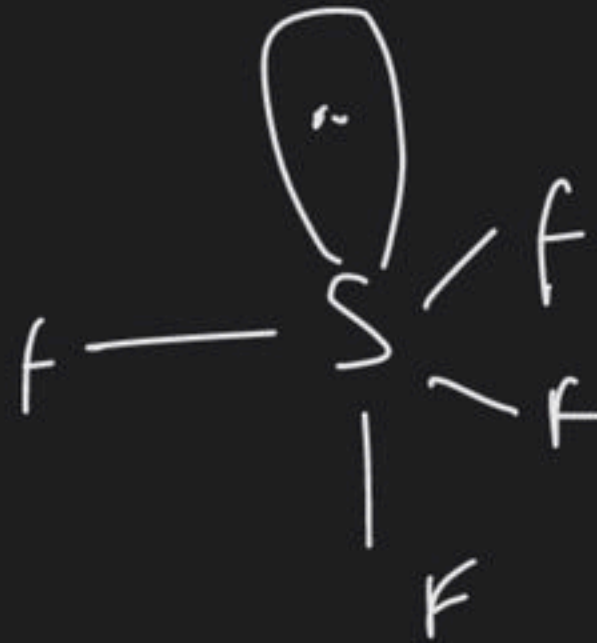
que

Which of the following Geometry of SF_4 is correct?

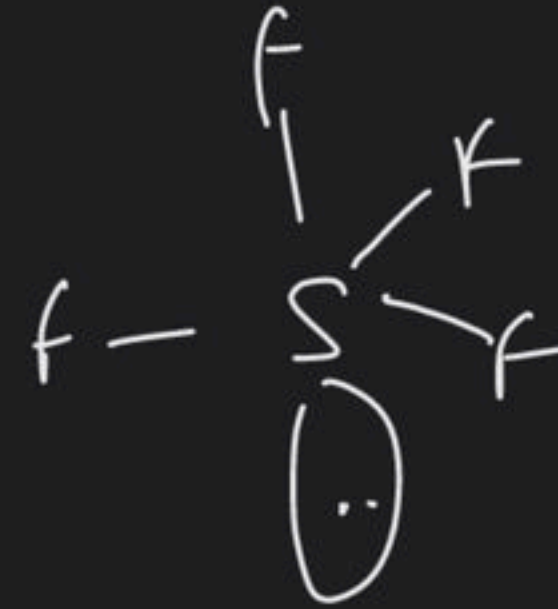
(1)



(2)



(3)



~~(4)~~ none of these



Question

from tavyan

Sir your most dedicated student





Question

from Anjali

Sir doubt

