



# History of Organic Chemistry, Bond Line Formula

Course on Nomenclature of Organic Compounds for Class XI

(\*) Basic level — Advance level

(\*) Doubt ✓

(\*) Theory Notes Register

(\*) Sheet ← Register

(\*) DRP ← Register

(\*) Book ← Register

(\*) Test paper

(\*) Test paper Analysis

(\*) O. chemistry (good)

(\*) KVPY (OC 11th copy)

(\*) PYQ (KVPY)

(\*) NSEC (PYQ)

ITT

12<sup>th</sup> class

1, 2, 3, 6, 8, 9, 10, 12

(Quiz add)

(\*) Theory Book



(<sup>11th</sup>) NICERT (after ZIT 11th conu)

BOOK

Problems & solution of

Organic Chemistry  
Cengage Publications  
(3e)

(Blue Book)  
BB

Suresh K. Mishra / SKM Siv)

Rank

# Organic chemistry

- ⇒ Retinal (organic compound)
- ⇒ Brain cells information transfer (organic chemistry)
- ⇒ Glucose (Energy source)

Willow bark (pain killer) (Aspirine)



# Chemistry

Organic chemistry

Inorganic chemistry

(\*) It is study of Compounds obtained from minerals & non living elements

(\*) It is study of  
Compounds obtained  
from living organisms/once  
living.

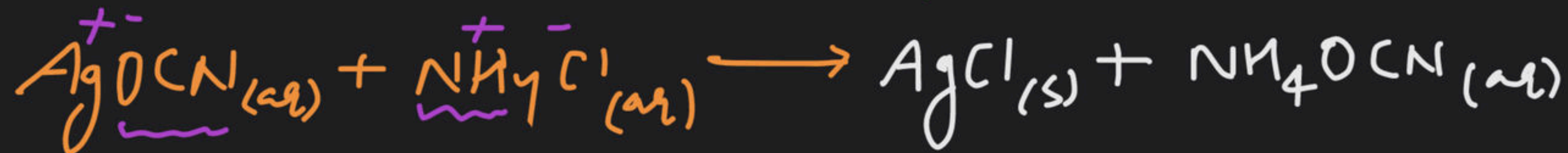


## Vital force theory

Organic Compounds are obtained in presence of a mysterious (vital force) & this theory is known as vital force theory.

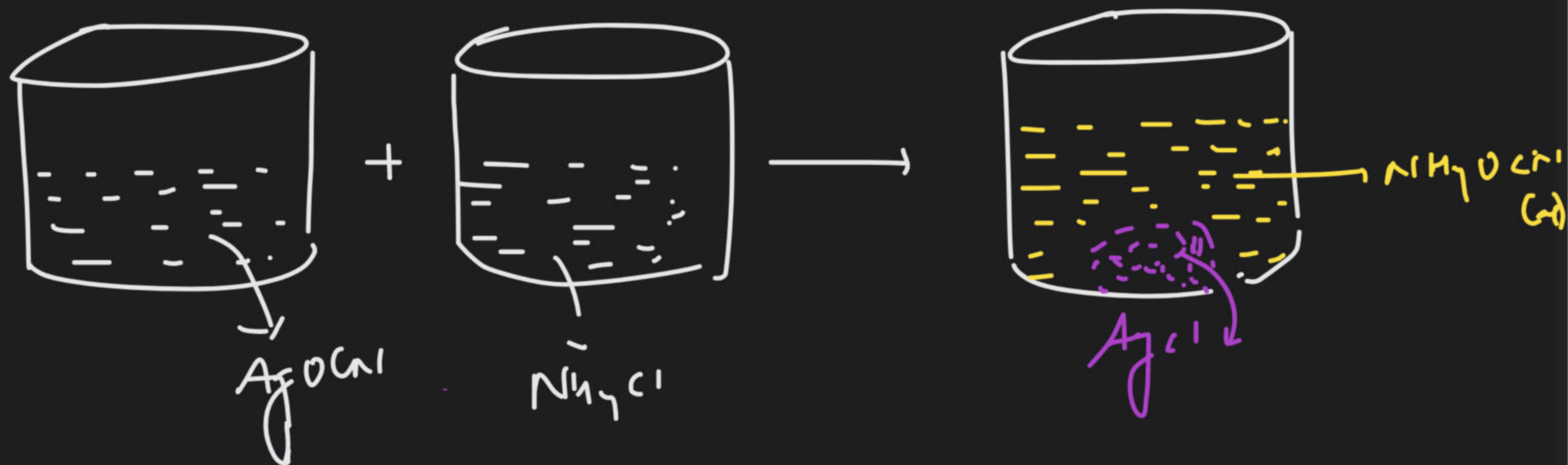
In 1828

Nohlar was trying to obtain Ammonium cyanate ( $\text{NH}_4\text{OCN}$ ) by reaction of aq. silver cyanate ( $\text{AgOCN}$ ) & Aq. Ammonium chloride ( $\text{NH}_4\text{Cl}$ )

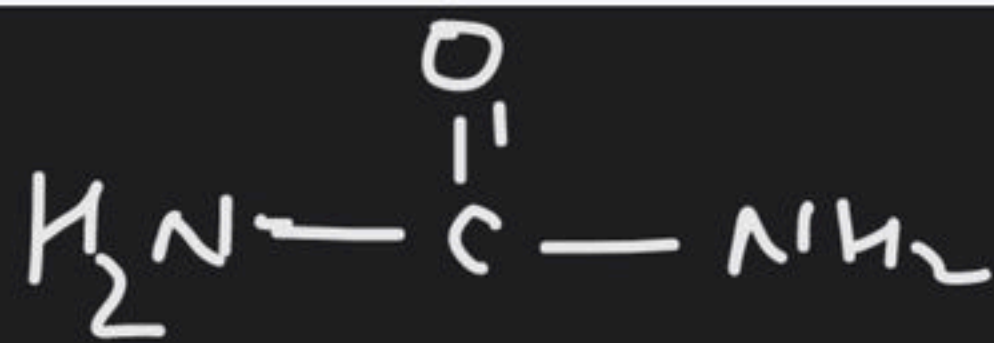
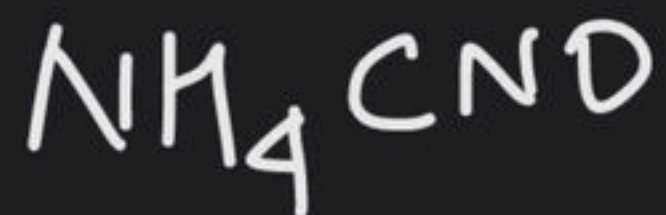




(a)



(\*) Carbamide (urea)  
(Organic Compound)



- (\*) Ammonium Cyanate
- (\*) Inorganic Compound.

- (x) urea
- (x) Carbamide
- (x) Organic Compound

- Note**
- (i) Organic Compound can be prepared in laboratory
  - (ii) Ammonium Cyanate & urea Both are isomers
  - (iii) discarded vital force theory.



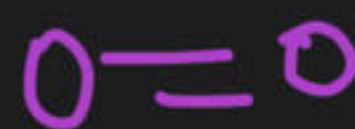
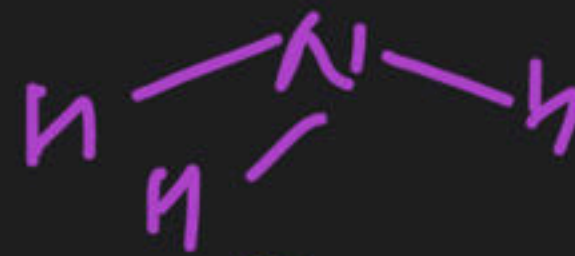
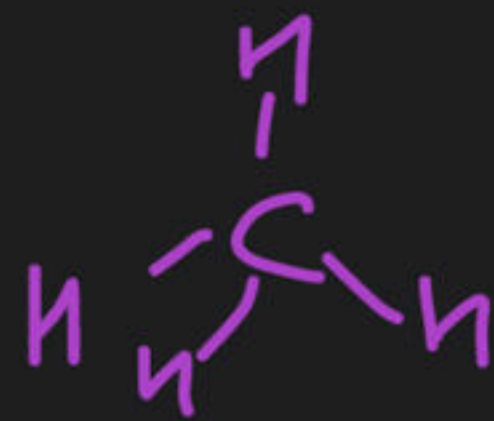
**Note** (i) <sup>IV</sup> Carbon Compounds

C  $\longrightarrow$  4 Bonds

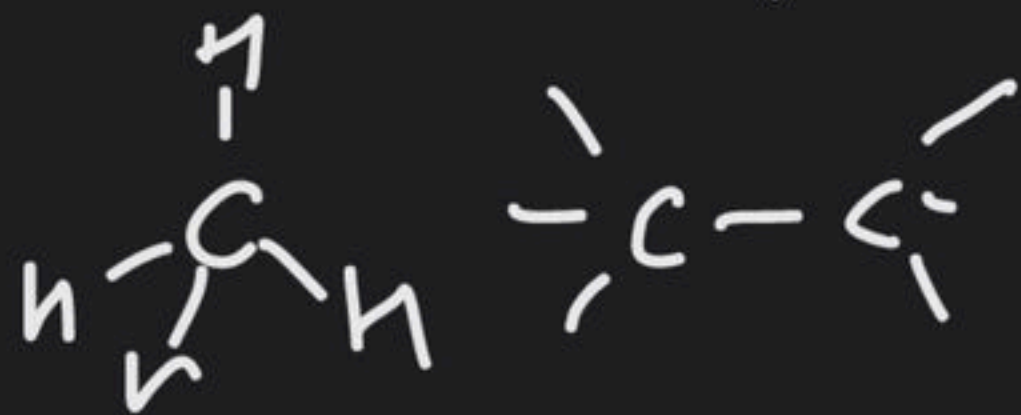
N  $\longrightarrow$  3 Bonds

O  $\longrightarrow$  2 Bonds

D/T/F/Cl/Br/I/H  $\longrightarrow$  1 Bond



(ii) Catenation ability of Carbon higher Catenation tendency of Carbon.



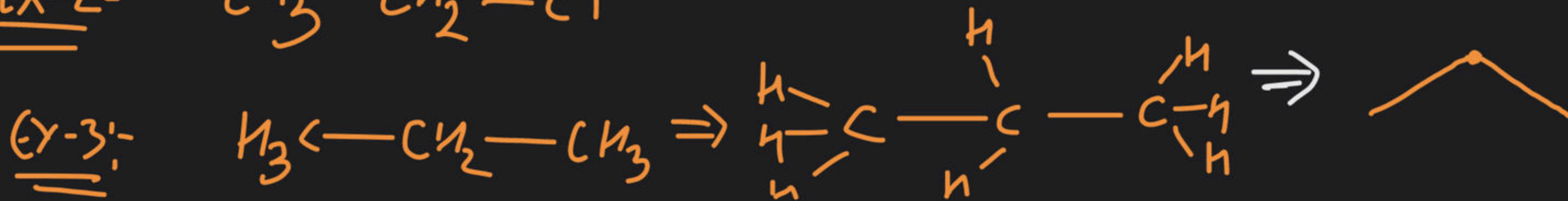
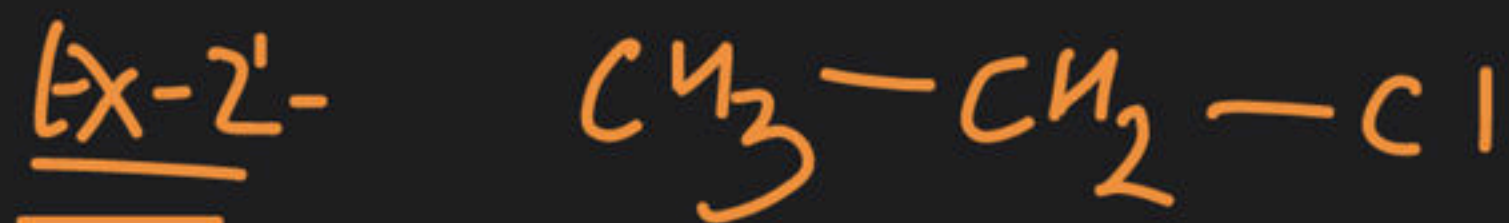
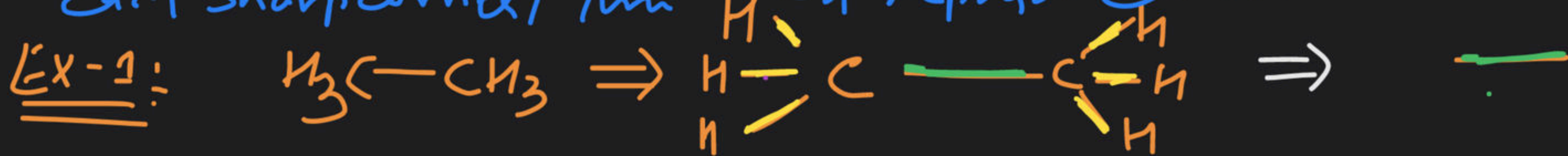
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# Bond Line Representation

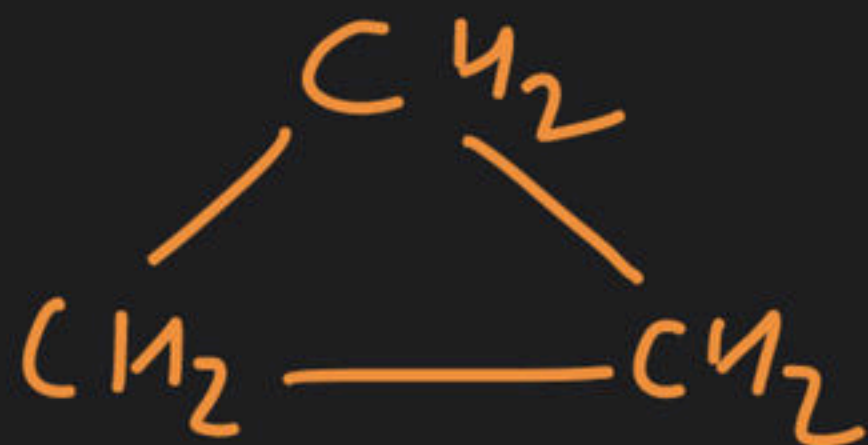
In Bond line formula

- (i) Never Draw C-H Bond
- (ii) Always Draw only C-C Bond as a line
- (iii) Rest all atoms other than C & H must be shown.  
Each sharp corner/Turn <sup>H</sup> <sub>and</sub> represent Carbon

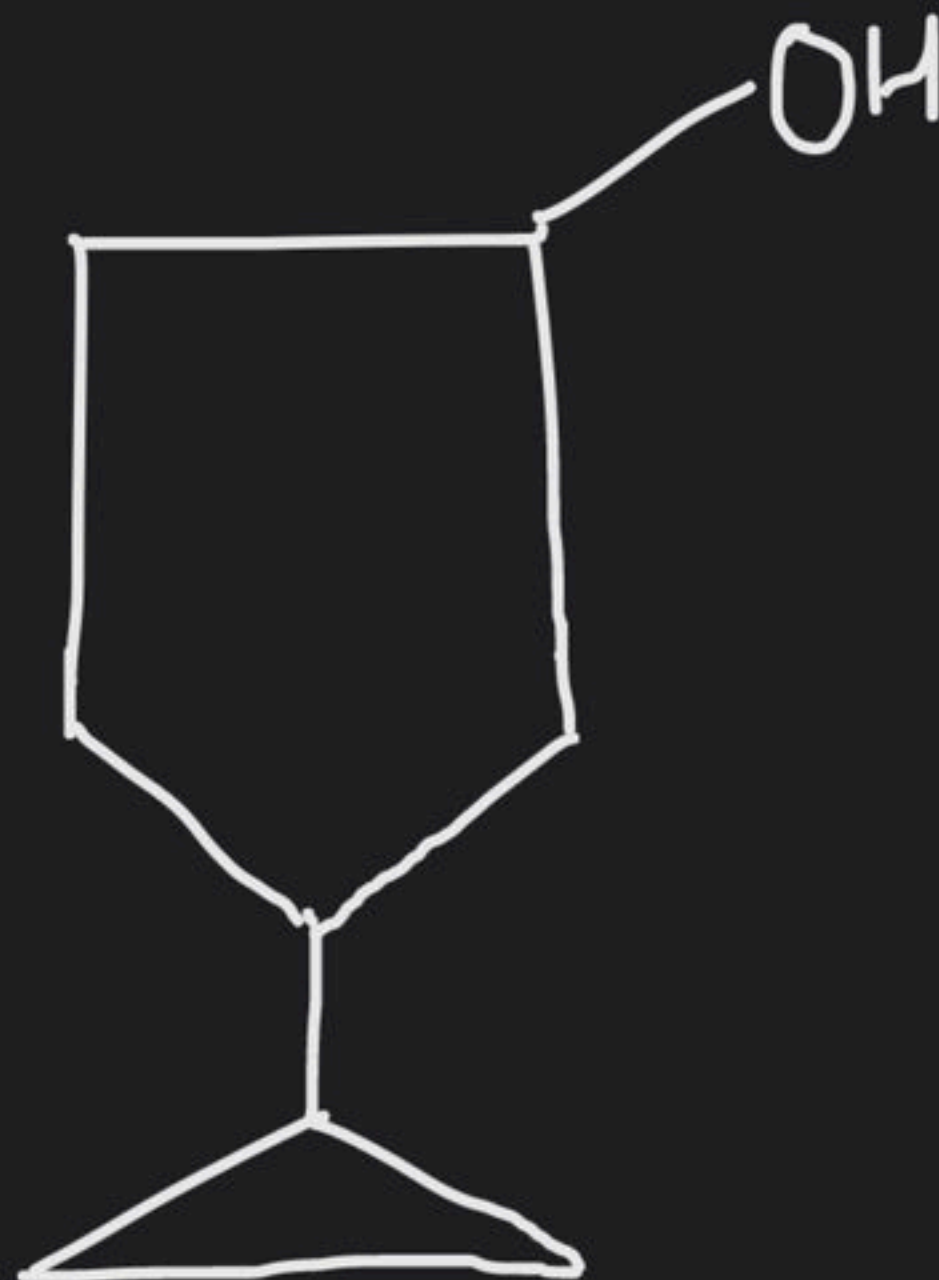
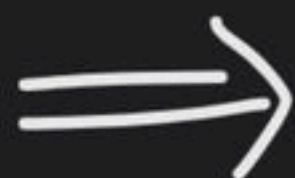
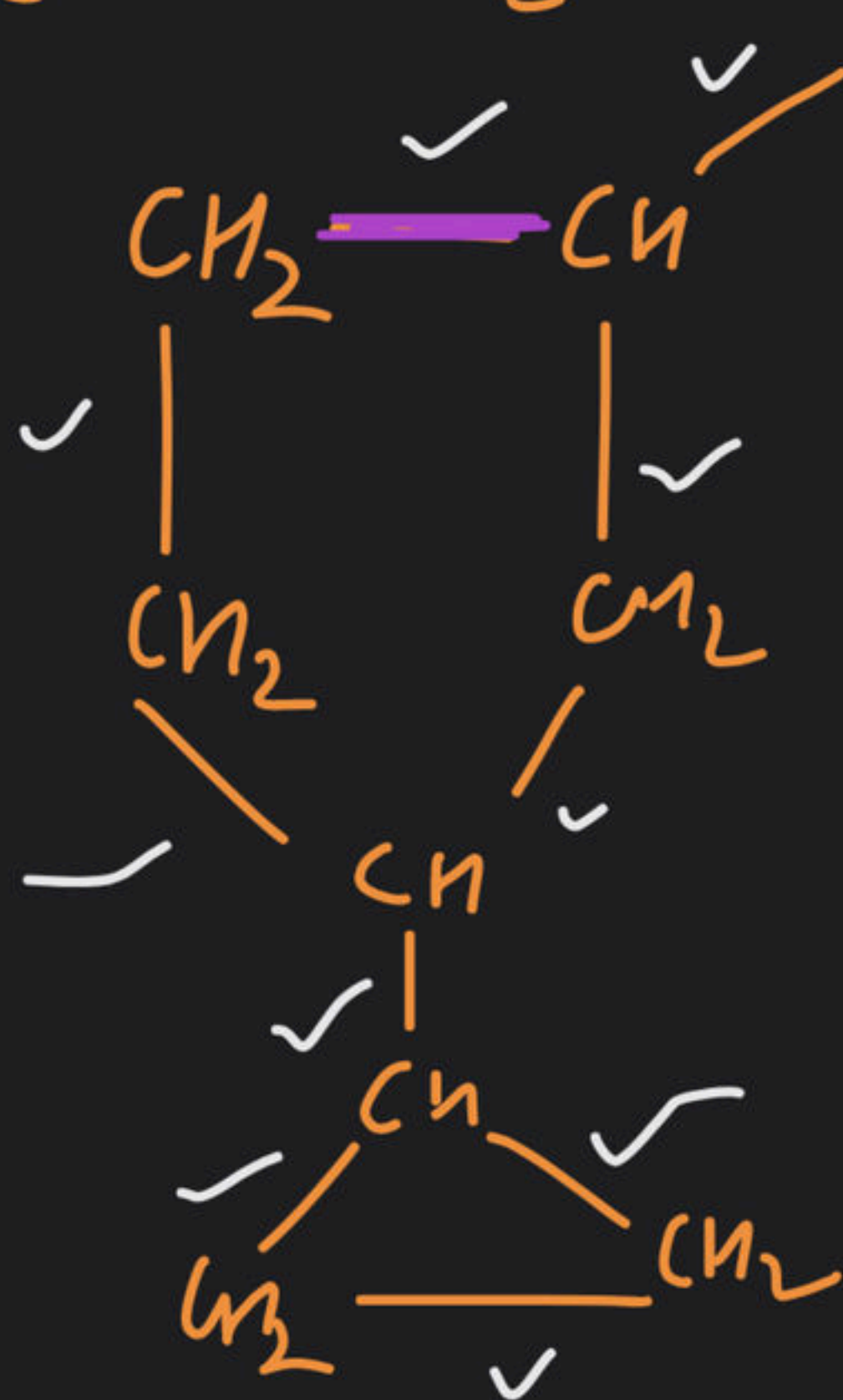




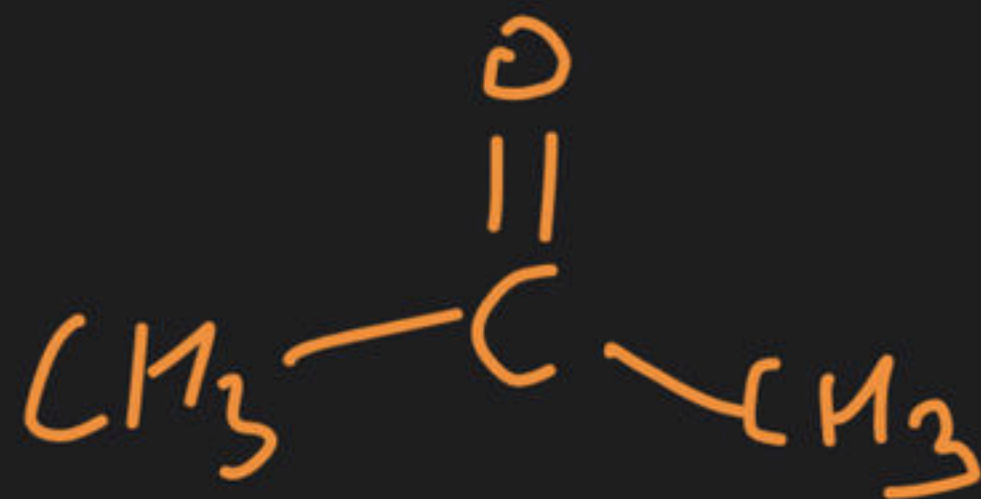
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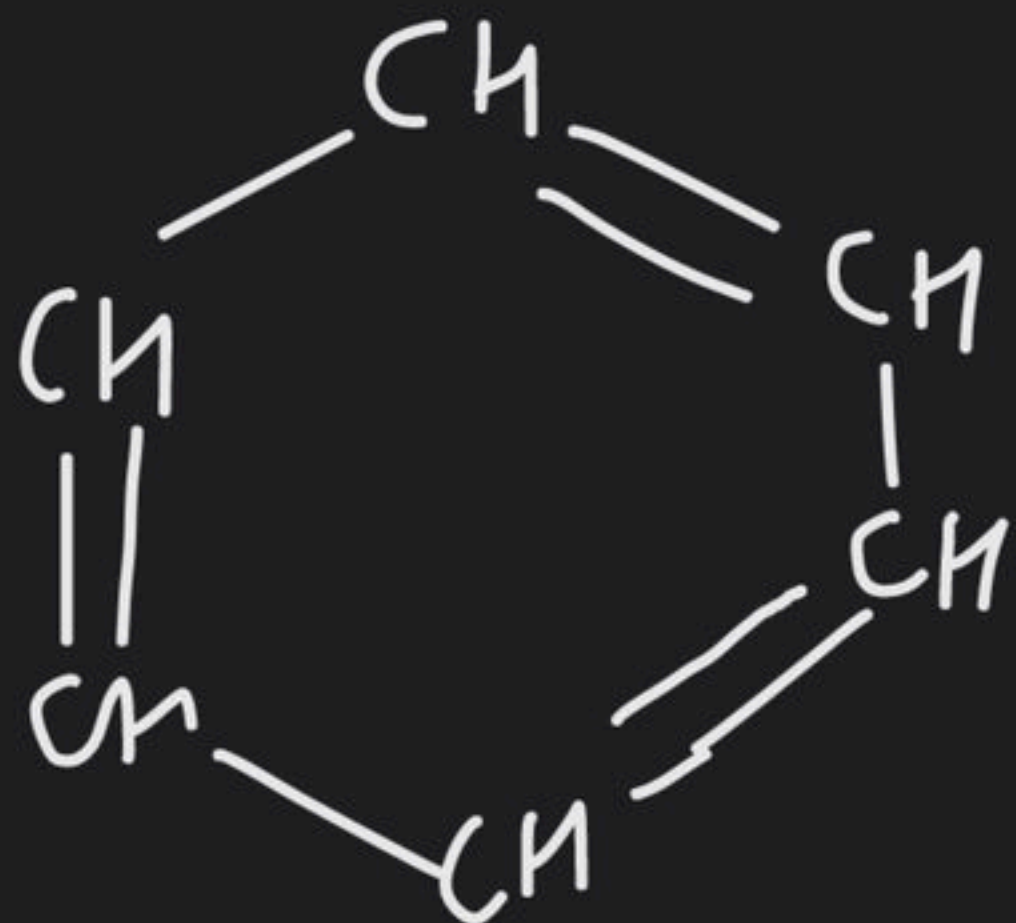
Ex-5:



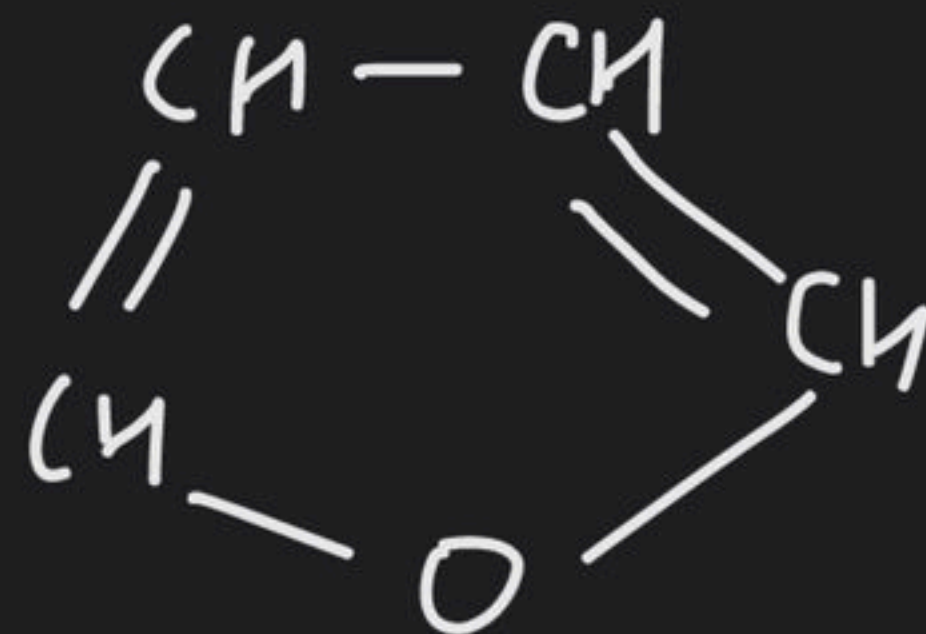
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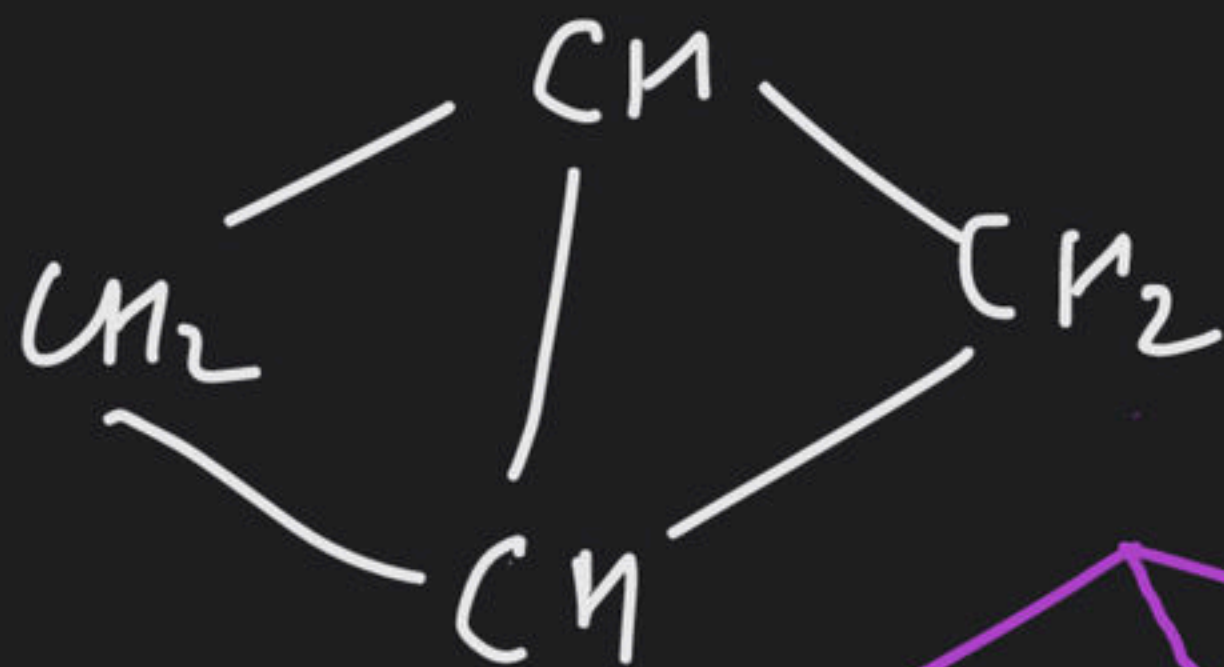
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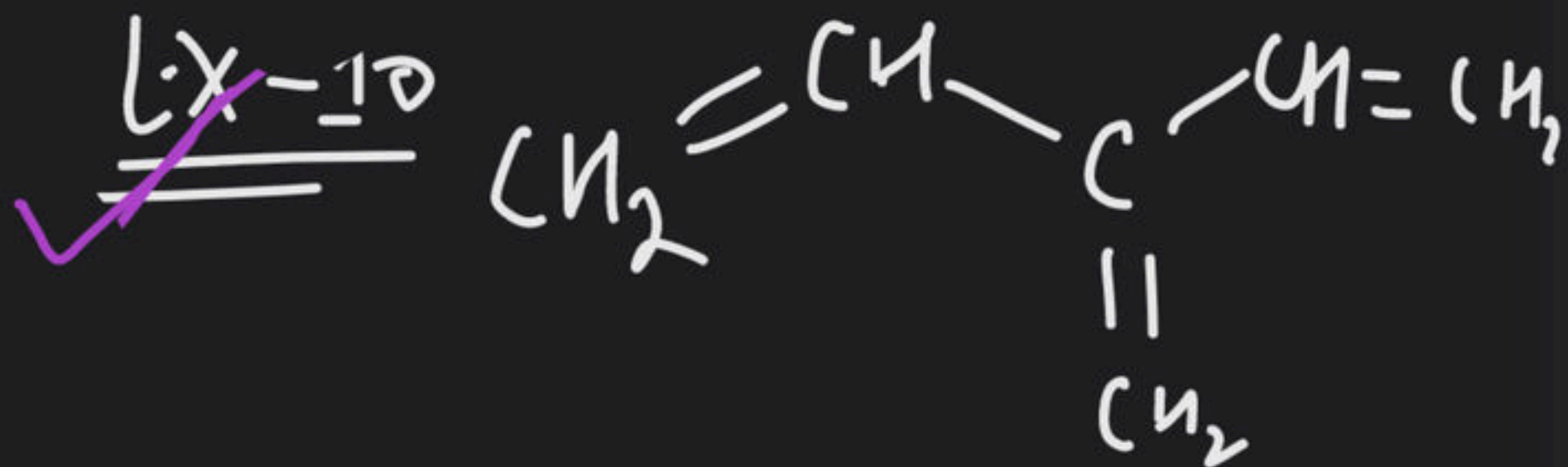
Ex-9:



Ex-8:

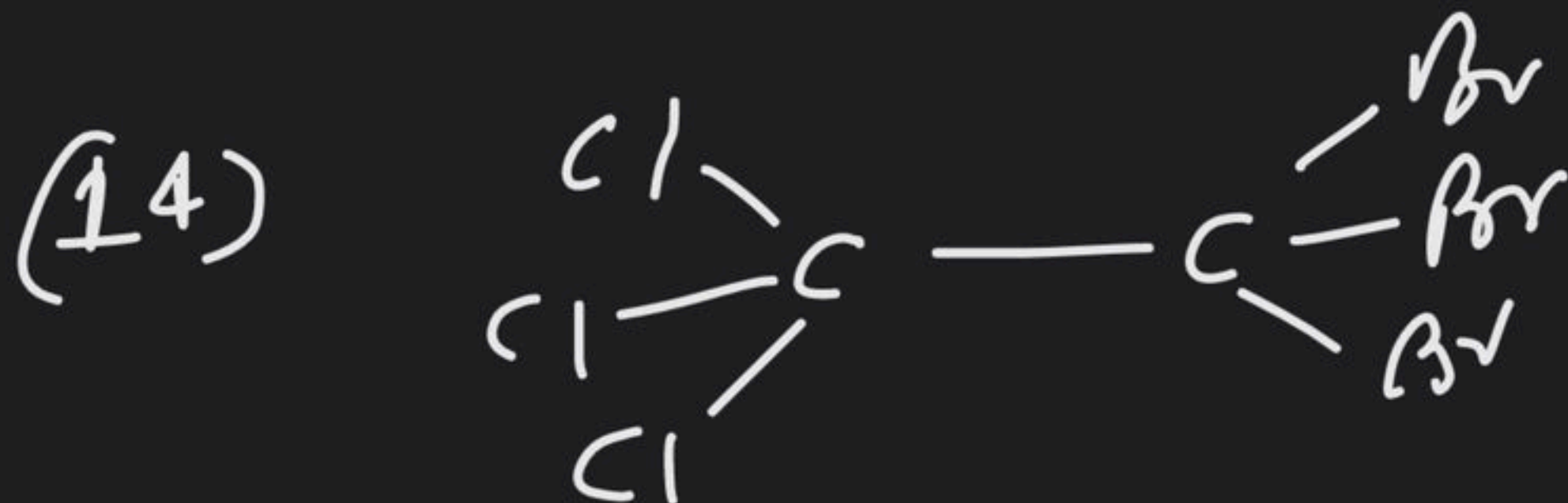
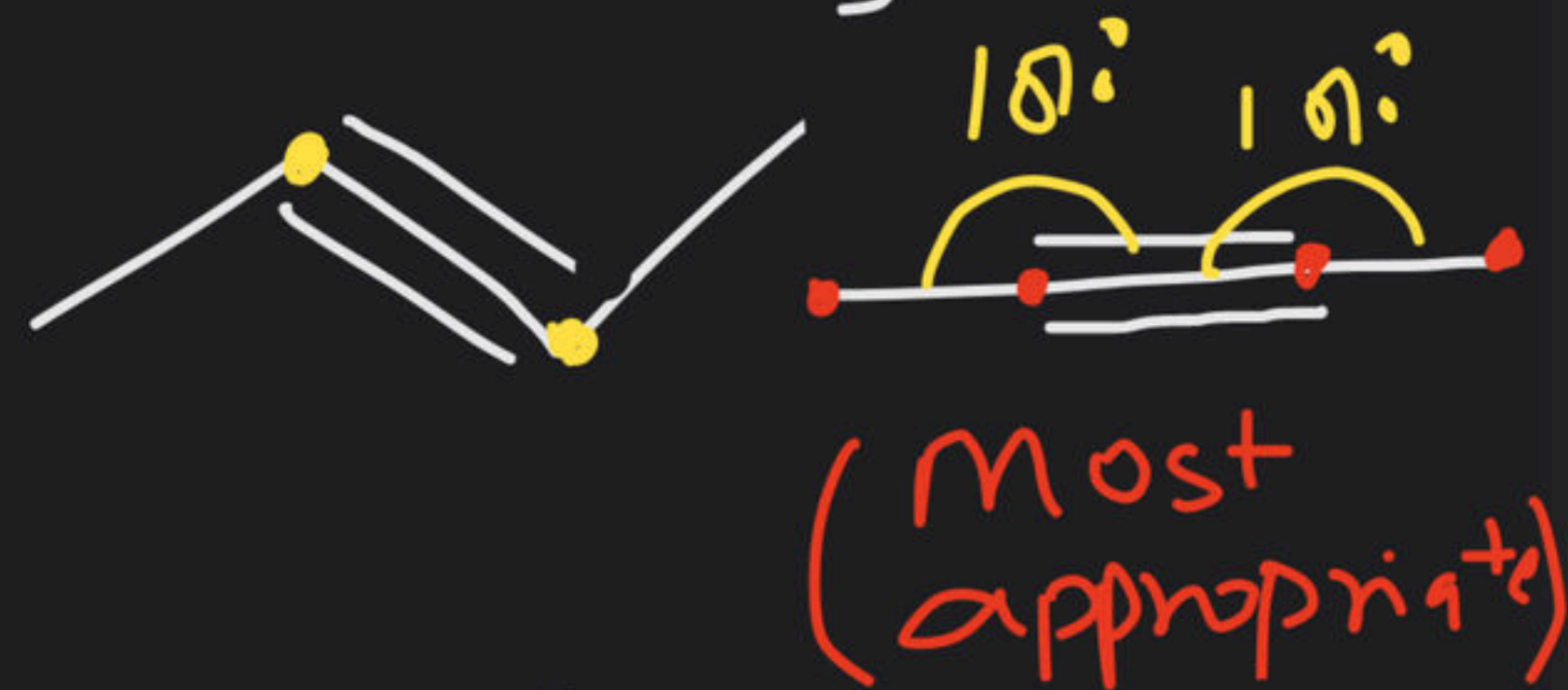
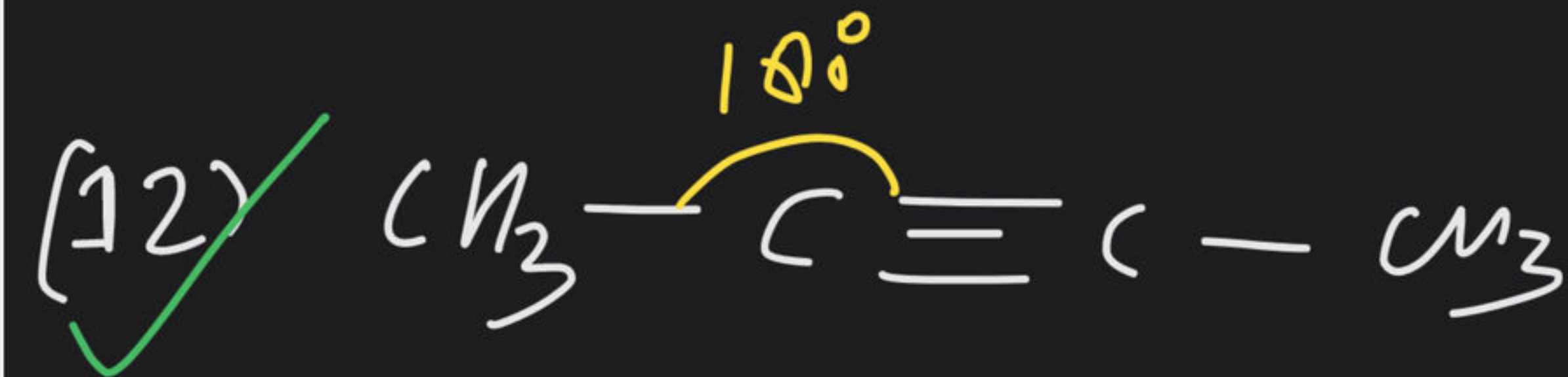
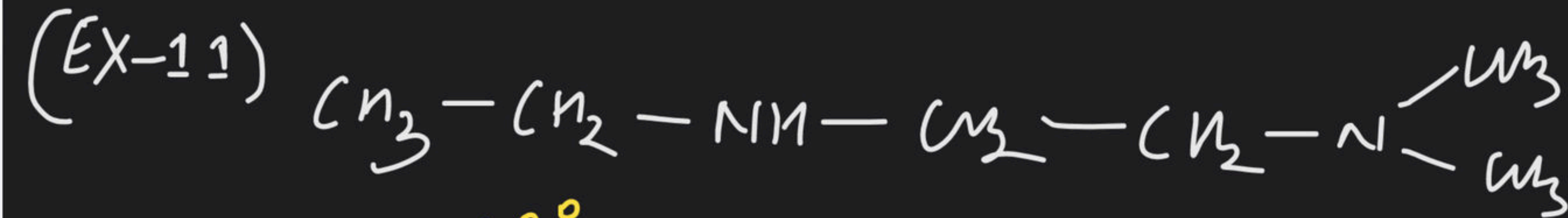


Ex-10

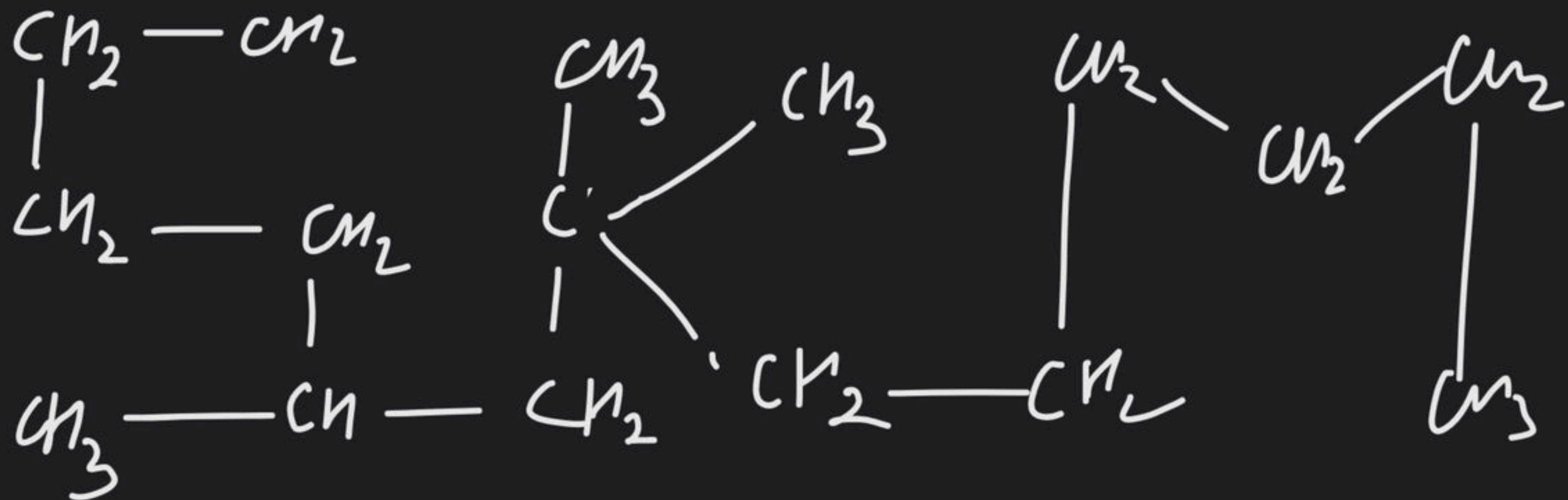




@SKM121



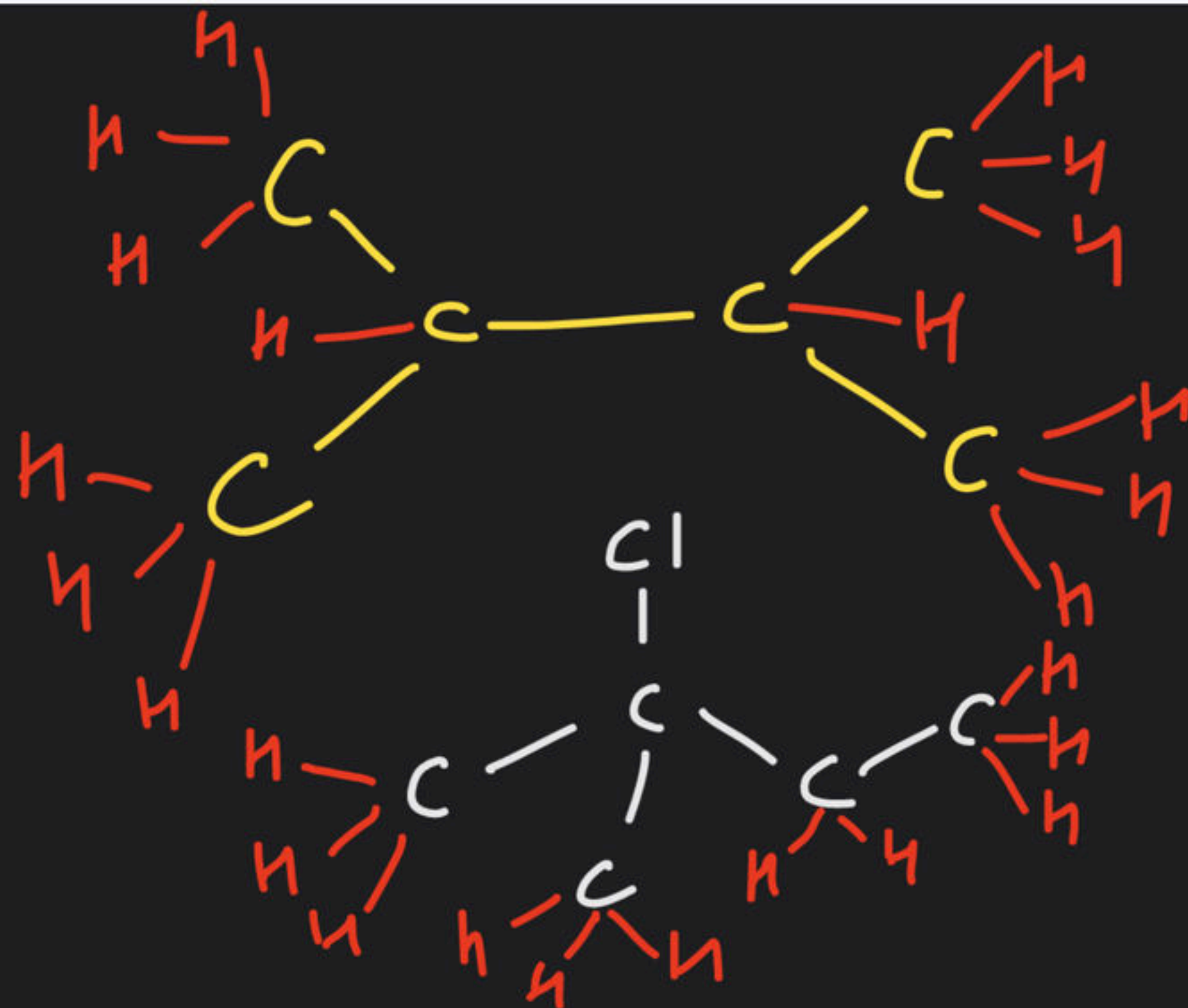
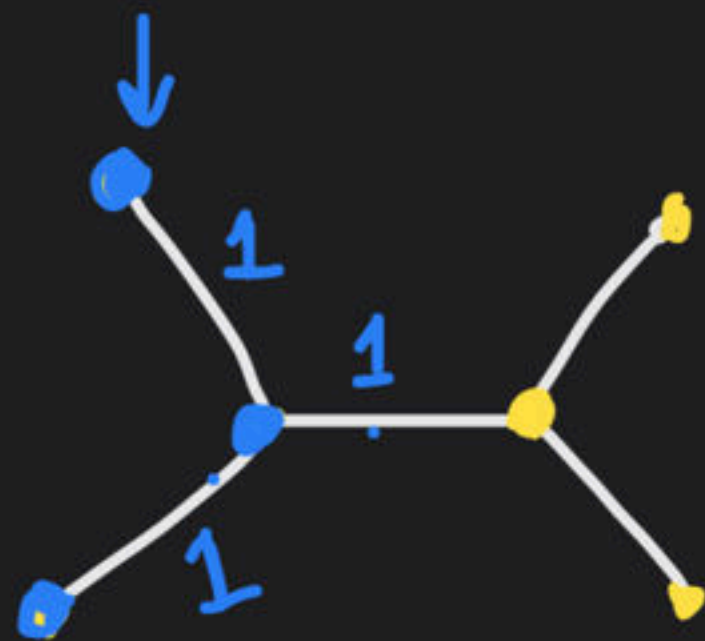
(15)



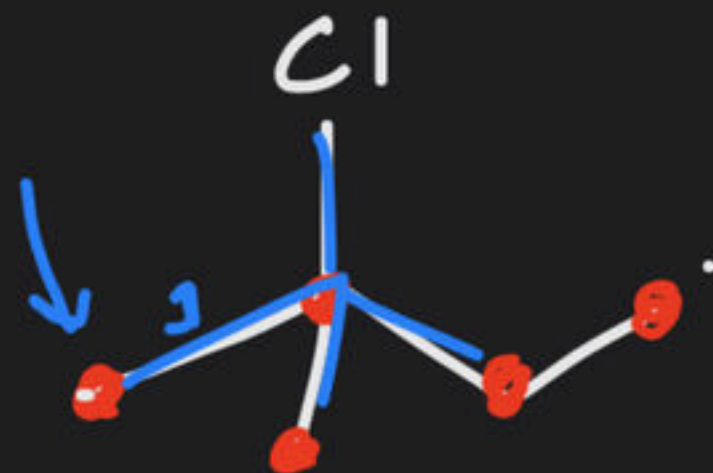
SKM



(16)

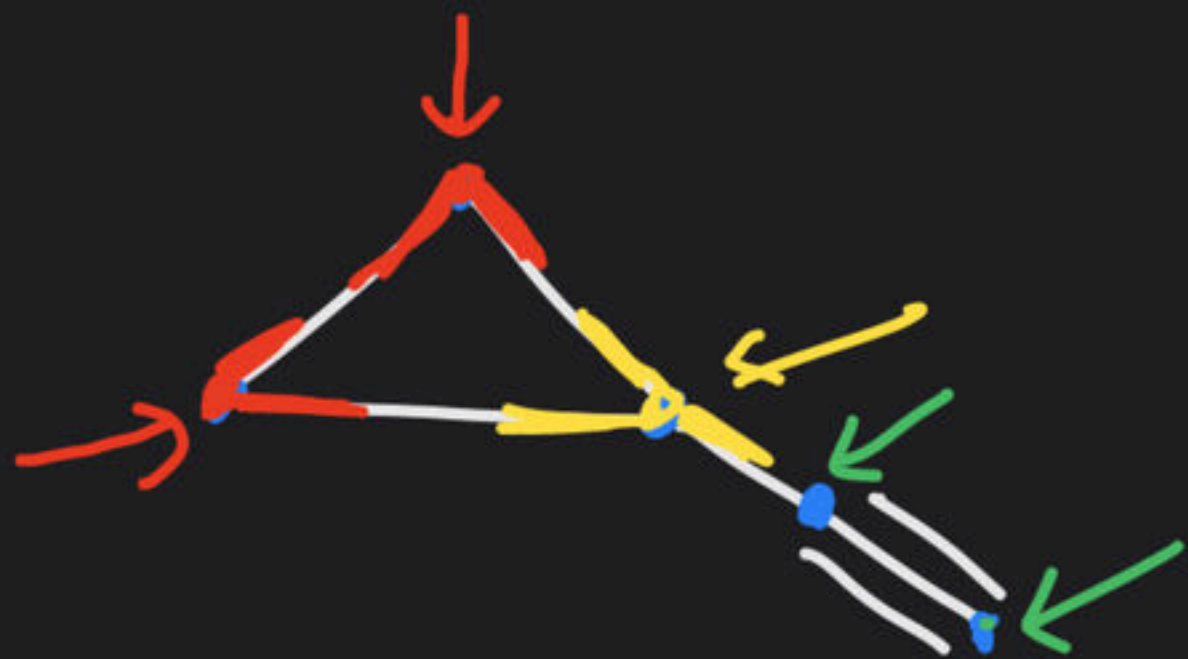


(17)



- ~~(A)~~  $C_3H_7Cl$  6%. ~~(B)~~  $C_5H_{11}Cl$   
~~(C)~~  $C_4H_{10}Cl$  34%. (D)  $C_5H_9Cl$

(18)



~~(A)~~

$C_4H_{10}$

(B)

$C_5H_{12}$

~~(C)~~

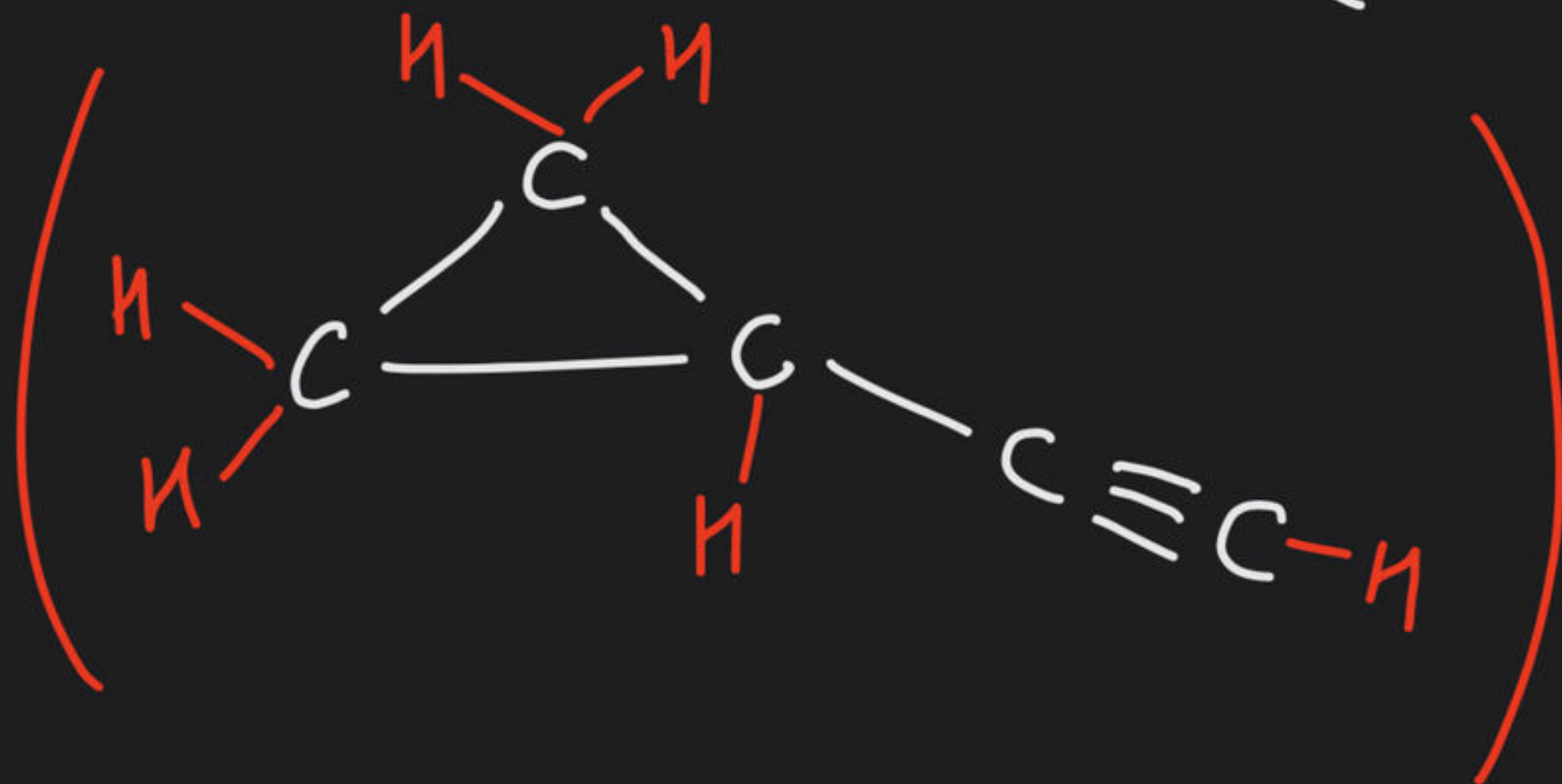
$C_6H_{14}$

~~(D)~~

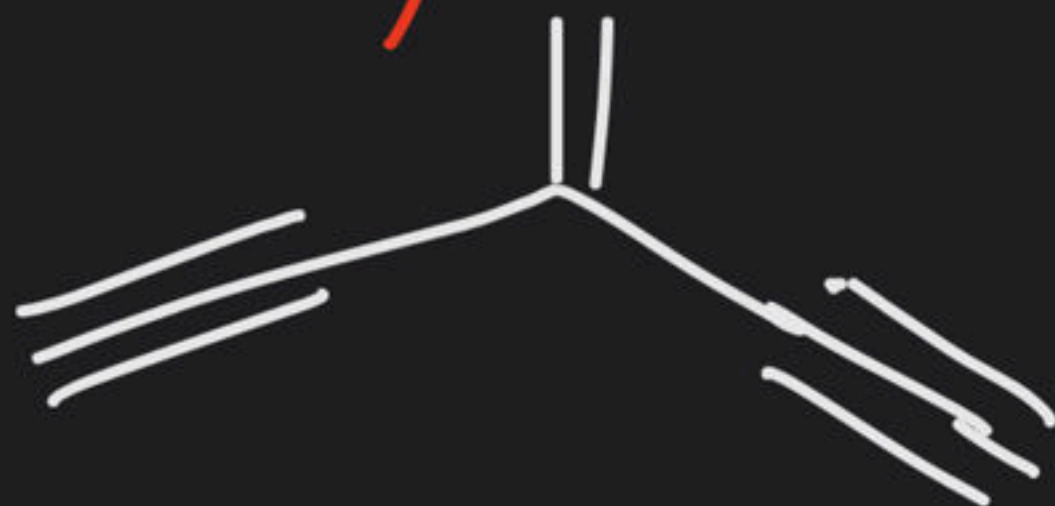
$C_4H_{10}$

(E)

$C_5H_{12}$

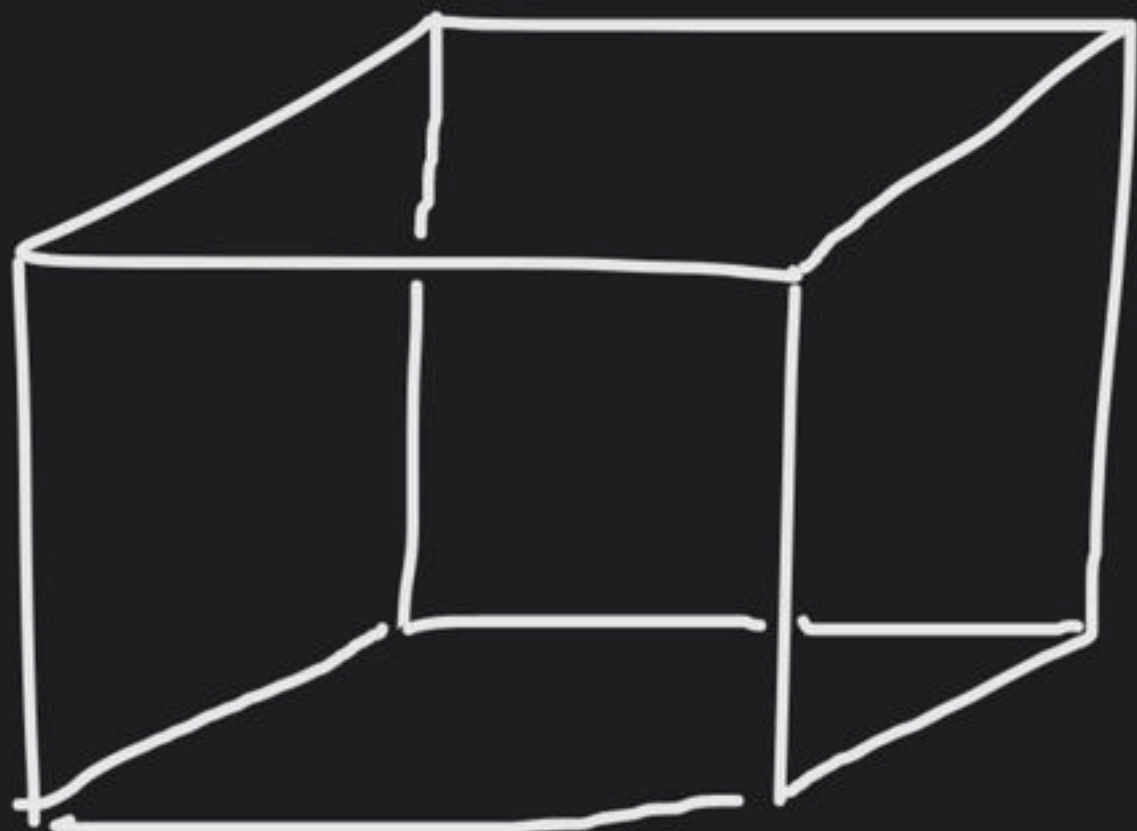


(19)

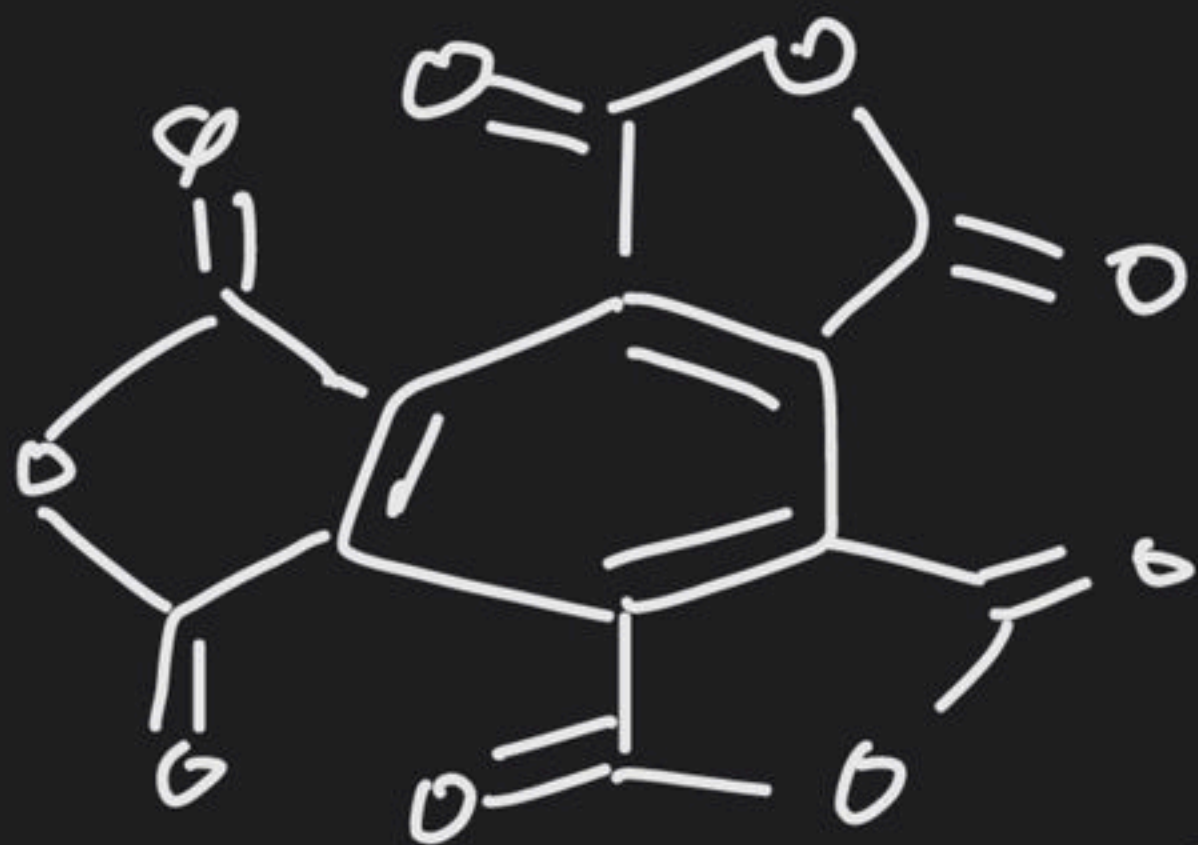




(20)



(21)

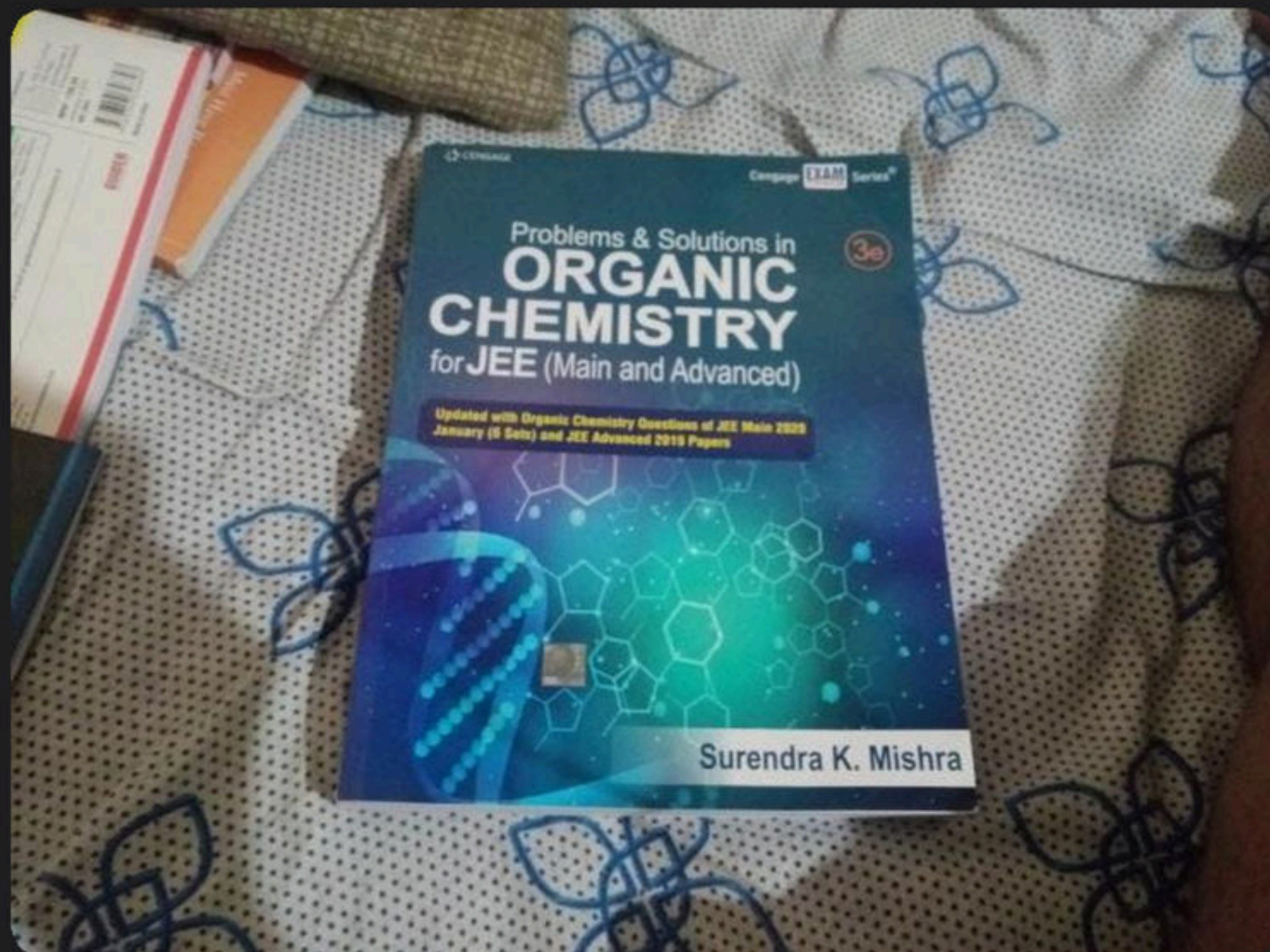




## Question

from Sumit rai

Apki book aagyi :)







## Question

from Aniket Sha 2024

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