

ARJUNA (NEET)

Classification of Elements & Periodicity in Properties

DPP-03

- Which of the following is best general electronic configuration of normal element?
 (1) $ns^{1-2} np^{0-6}$ (2) $ns^{1-2} np^{1-5}$
 (3) $ns^{1-2} np^{0-5}$ (4) $ns^{1-2} np^{1-6}$
- Which of the following elements belong to alkali metals ?
 (1) $1s^2, 2s^2 2p^2$
 (2) $1s^2, 2s^2 2p^6, 3s^2 3p^6 3d^{10}, 4s^2 4p^6, 5s^1$
 (3) $1s^2, 2s^2 2p^5$
 (4) None of these
- Which of the following statement is wrong?
 (1) Total no. of liquid elements in the periodic table.....Six
 (2) First metal element in the periodic table is....Li
 (3) All type of elements are present in 6th period
 (4) Iodine is a gaseous element.
- The IUPAC name of the element which is placed after Db_{105} in the periodic table, will be :-
 (1) Un nil pentium (2) Un un nilium
 (3) Un nil hexium (4) Un nil quadium
- The element with atomic number $Z=118$ will be :-
 (1) Noble gas
 (2) Transition metal
 (3) Alkali metal
 (4) Alkaline earth metal
- The electronic configuration of d-block elements is exhibited by :-
 (1) $ns^{1-2}(n-1)d^{1-10}$ (2) $ns^2(n-1)d^{10}$
 (3) $(n-1)d^{10}s^2$ (4) ns^2np^5
- If the atomic number of an element is 33, it will be placed in the periodic table in the
 (A) 1st group (B) 3rd group
 (C) 15th group (D) 17th group
- An element has 56 nucleons in nucleus and if it is isotonic with ${}_{30}Y^{60}$. Which group and period does it belong to?
 (A) 8th group, 4th period
 (B) 14th group, 3rd period
 (C) 12th group, 3rd period
 (D) 12th group, 4th period
- An element has electronic configuration $[Xe] 4f^7, 5d^1, 6s^2$. It belongs to block of the periodic table.
 (A) s (B) p
 (C) d (D) f
- Which of the following elements do not belong to the family indicated?
 (A) Cu – Coinage metal
 (B) Ba – Alkaline earth metal
 (C) Zn – Alkaline earth metal
 (D) Xe – Noble gas

ANSWER KEY

1. (C)
2. (B)
3. (D)
4. (C)
5. (A)
6. (A)
7. (C)
8. (A)
9. (D)
10. (C)



Note - If you have any query/issue



Mail us at support@physicswallah.org
