

ARJUNA (NEET)

Some Basic Concepts of Chemistry

DPP-04

- A substance, on analysis, gave the following percentage composition : Na = 43.4%, C = 11.3%, O = 45.3%. Calculate the empirical formula. (Na = 23, C = 12, O = 16)

(A) NaCO_3 (B) NaO
(C) Na_2CO_3 (D) CO_2
- An organic compound on analysis gave the following data : C = 57.82%, H = 3.6% and 38.58% is oxygen. Find its empirical formula.

(a) $\text{C}_4\text{H}_3\text{O}_2$ (B) CH_2O
(C) CH_4O (D) $\text{C}_3\text{H}_4\text{O}_2$
- Two elements X and Y (atomic mass of X = 75, Y = 16) combine to give a compound having 76% of X. The formula of the compound is

(A) XY (B) X_2Y
(C) X_2Y_2 (D) X_2Y_3
- The percentage of nitrogen in HNO_3 is

(A) 22.22% (B) 35%
(C) 28.57% (D) 45%
- What is the ratio of empirical formula mass to molecular formula mass of benzene?

(A) 1 : 6 (B) 2 : 3
(C) 6 : 1 (D) 3 : 2
- The compound in which mass percentage of carbon is 75% and that of hydrogen is 25% is

(A) C_2H_6 (B) C_2H_2
(C) CH_4 (D) C_2H_4
- Identify the molecule having empirical formula CH_2O

(A) CH_3CHO
(B) $\text{HO}-\text{CH}_2-\text{CH}_2-\text{OH}$
(C) $\begin{array}{c} \text{COOH} \\ | \\ \text{COOH} \end{array}$
(D) CH_3COOH

ANSWERS

1. (C)
2. (A)
3. (D)
4. (A)
5. (A)
6. (C)
7. (D)



Note - If you have any query/issue

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