

ARJUNA (NEET)

Classification of Elements & Periodicity in Properties

DPP-04

1. Arrange the elements in increasing order of atomic radius Na, Rb, K, Mg:
 (A) Na, K, Mg, Rb
 (B) K, Na, Mg, Rb
 (C) Mg, Na, K, Rb
 (D) Rb, K, Mg, Na
2. Which of the following pairs of elements have almost similar atomic radii?
 (A) Zr, Hf (B) Mo, W
 (C) Co, Ni (D) All
3. The screening effect of d- electrons is :-
 (A) Equal to the p - electrons
 (B) Much more than p - electrons
 (C) Same as f - electrons
 (D) Less than p - electrons
4. Which of the following is not isoelectronic series?
 (A) Cl^- , P^{3-} , Ar (B) N^{3-} , Ne, Mg^{+2}
 (C) B^{+3} , He, Li^+ (D) N^{3-} , S^{2-} , Cl^-
5. Which group of atoms have nearly same atomic radius?
 (A) Na, K, Rb, Cs (B) Li, Be, B, C
 (C) Fe, Co, Ni (D) F, Cl, Br, I
6. Atomic radii of Fluorine and Neon in Angstrom units are given by:-
 (A) 0.72, 1.60 (B) 1.60, 1.60
 (C) 0.72, 0.72 (D) None of these
7. Which of the following has largest radius?
 (A) $1s^2 2s^2 2p^6 3s^2$
 (B) $1s^2 2s^2 2p^6 3s^2 3p^1$
 (C) $1s^2 2s^2 2p^6 3s^2 3p^3$
 (D) $1s^2 2s^2 2p^6 3s^2 3p^5$
8. Which of the following order of atomic/ionic radius is not correct?
 (A) $\text{I}^- > \text{I} > \text{I}^+$
 (B) $\text{Mg}^{+2} > \text{Na}^+ > \text{F}^-$
 (C) $\text{P}^{+5} < \text{P}^{+3}$
 (D) $\text{Li} > \text{Be} > \text{B}$
9. Correct order of ionic radii is
 (A) $\text{Ti}^{4+} < \text{Mn}^{7+}$
 (B) $^{37}\text{Cl}^- < ^{35}\text{Cl}^-$
 (C) $\text{K}^+ > \text{Cl}^-$
 (D) $\text{P}^{3+} > \text{P}^{5+}$
10. In an anion :-
 (A) Number of proton decreases
 (B) Protons are more than electrons
 (C) Effective nuclear charge is more
 (D) Radius is larger than neutral atom

ANSWER KEY

1. (C)
2. (D)
3. (D)
4. (D)
5. (C)
6. (A)
7. (A)
8. (B)
9. (D)
10. (D)



Note - If you have any query/issue



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