

Course on States of Matter for Class XI

$$\frac{3}{293}$$

$$\frac{313}{293}$$

$$\frac{3}{293} + 3H_2 - 3B + 6HQ.$$

$$\frac{702}{502} = \times_{02}$$

Maxwell distribution of speed of Molecules Speed of a molecules changes continuously due to collision but total KE remains Constant truerefore it can be Concluded that will purticles in given speed range remain construt with filme.

 $\frac{dN}{dN} = \frac{4\pi}{2\pi RT} \left( \frac{M}{2\pi RT} \right)^{3/2} \ln^2 \frac{M^2}{2\pi RT} dm$ dN = nv. of particle having speed bett uto utdu

N = total no of particle

M = Molman dN

 $J = 3 \chi^3 e^{-\chi^2}$ 

dN, dN, dNz + dN2 x42 dN, XY, Javy -4N1 + dN2 - --. = JSRT TM Jarg

4N1 X 42 + dN2 x 42 - - -( dNXh2)/2

Vrms - J3RT M find Vrms spread of H2(j) at 300 K.

$$= \frac{3 \times 25 \times 300}{2 \times 16^{3}}$$

representation of egN 12x well U to 4+1 Imps

1) Area =  $\int \frac{dN}{N} \cdot dx = \int \frac{dN}{N}$ - Fraction of particles

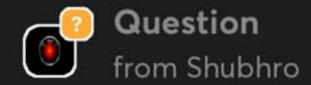
2) Total area = 1 (3) Vmps (most probable spred)

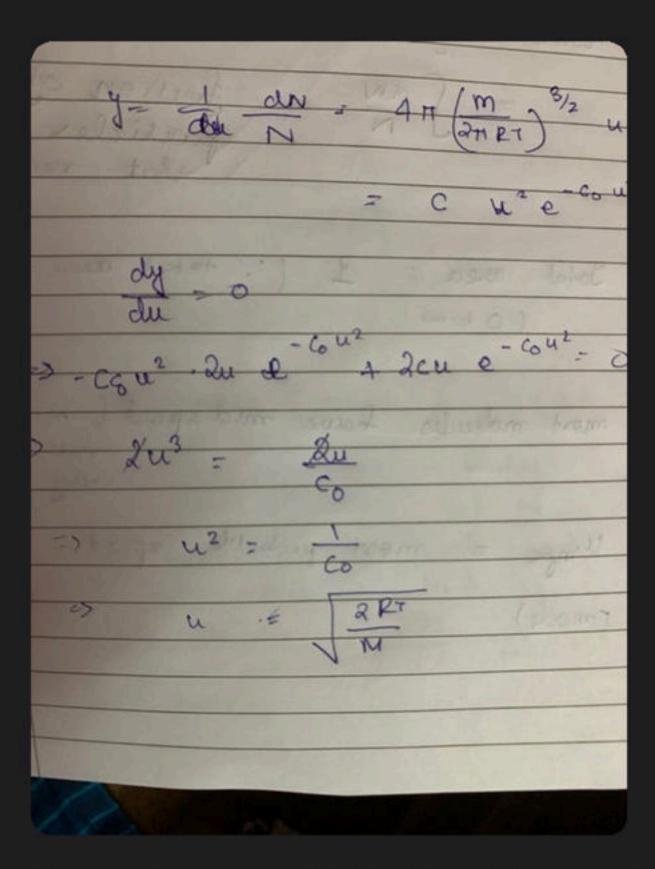
$$\frac{1}{du}\frac{dN}{N} = y = 4\pi \left(\frac{M}{2\pi RT}\right)^{3/2}u^{2}e^{-Mu^{2}/2RT}$$

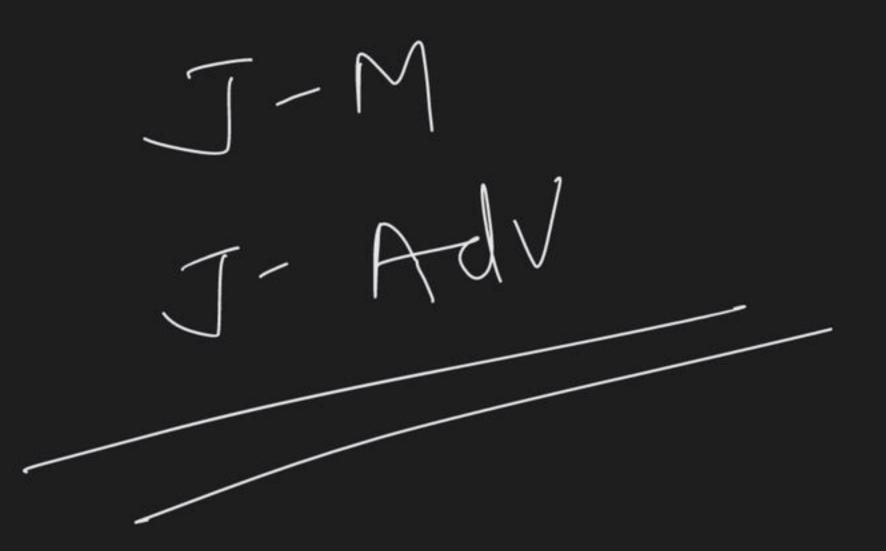
$$y = Cu^{2}e^{-Mu^{2}/2RT}$$

$$V_{mps} = \sqrt{2RT}$$

$$M$$







du du

Areg- ( d/dN x dh

dN

Area - 1 7dx

 $\frac{\lambda}{\lambda} = \frac{\lambda^{2}}{2}$   $\frac{\lambda}{\lambda} = \frac{\lambda^{2}}{2} + \lambda e^{-\lambda^{2}} \left(-2\lambda\right)$