CHEMISTRY

NEET

CLASS-XI

MODULE-01

Some Basic Concept of Chemistry

| Structure of Atom | Classification of Elements
| Chemical Bonding & Molecular Structure

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Physics Wallah

Topic-wise Questions

Uncertainity in Measurement

1. The number of significant figures in 0.0045 are
   a. Two
   b. Three
   c. Four
   d. Five.

2. Light travels with a speed of 3×10^8 m/sec. The distance travelled by light in 1 Femto sec is:
   a. 0.03 mm
   b. 0.003 mm
   c. 3 mm
   d. 0.0003 mm

3. Area of nuclear cross-section is measured in "Barn". It is equal to:
   a. $10^{-20} \, m^2$
   b. $10^{-30} \, m^2$
   c. $10^{-28} \, m^2$
   d. $10^{-14} \, m^2$

4. Two students X and Y report the mass of the same substance as 7.0 g and 7.00 g respectively, which of the following statement is correct?
   a. Both are equally accurate
   b. X is more accurate than Y
   c. Y is more accurate than X
   d. Both are inaccurate scientifically

5. The number of significant figures in value of π are:
   a. 1
   b. 2
   c. 3
   d. ∞

6. 5.041 has how many significant figures.
   a. 1
   b. 2
   c. 3
   d. 4

7. The correctly reported answer of the addition of 29.4406, 3.2 and 2.25 will have significant figures:
   a. 3
   b. 4
   c. 2
   d. 5

8. What is the area of rectangle which is 12.34 cm wide and 1.23 cm long?
   a. $15.2 \, m^2$
   b. $15.19 \, cm^2$
   c. $16.2 \, cm^2$
   d. $16.2 \, m^2$

9. If an object has a mass of 0.2876 g, then find the mass of nine such objects:
   a. 2.5884 g
   b. 2.5886 g
   c. 2.588 g
   d. 2.5 g

10. The value of Plank's constant is 6.62618×10^{-34} Js. The number of significant figures in it is
    a. Six
    b. Five
    c. Three
    d. Thirty four

Law of Chemical Combinations

11. In Habers process, the volume at S.T.P of ammonia relative to the total volume of reactants at STP is :
    a. One fourth
    b. One half
    c. Same
    d. Three fourth

12. 6 g of carbon combines with 32 g of sulphur to form CS_2, 12 g of C also combine with 32 g oxygen to form CO_2. 10 g of sulphur combines with 10 g of oxygen to form Sulphur dioxide. Which law is illustrated by this?
    a. Law of multiple proportions
    b. Law of constant composition
    c. Law of reciprocal proportions
    d. Gay Lussac's law

13. Which of the following data illustrates the law of conservation of mass?
    a. 56 g of C reacts with 32 g of Oxygen to produce 44 g of CO_2
    b. 1.70 g of $AgNO_3$ reacts with 100 ml of 0.1M HCl to produce 1.435 g of AgCl and 0.63 g of HNO_3
    c. 12 g of C is heated in vacuum and on cooling, there is no change in mass
    d. 36 g of S reacts with 16 g of O_2 to produce 48 g of SO_2

14. One part of an element A combines with two parts of another element B, 6 parts of element C combines with 4 parts of B. If A and C combine together the ratio of their weights, will be governed by
    a. law of definite proportion
    b. law of multiple proportion
    c. law of reciprocal proportion
    d. law of conservation of mass

15. The law of conservation of mass holds good for all of the following except.
    a. All chemical reactions
    b. Nuclear reaction
    c. Endothermic reactions
    d. Exothermic reactions

16. The % of copper and oxygen in samples of CuO obtained by different methods were found to be the same. This proves the law of:
    a. Constant Proportion
    b. Reciprocal Proportion
    c. Multiple Proportion
    d. Conservation of mass.

17. Two elements X and Y combine in gaseous state to form XY in the ratio 1:35.5 by mass. The mass of Y that will be required to react with 2 g of X is:

a. 7.1 g
b. 3.55 g
c. 71 g
d. 35.5 g

18. 4.4 g of an oxide of nitrogen gives 2.24 L of nitrogen and 60 g of another oxide of nitrogen gives 22.4 L of nitrogen at S.T.P. The data illustrates:

a. Law of conservation of mass
b. Law of constant proportions
c. Law of multiple proportions
d. Law of reciprocal proportions

19. "The total mass of reactants is always equal to the total mass of products in a chemical reaction." This statement is konwn as

a. Law of conservation of mass
b. Law of definite proportions
c. Law of equivalent weights
d. Law of combining masses

20. The law of multiple proportions is illustrated by the two compounds

a. Sodium chloride and sodium bromide
b. Ordinary water and heavy water
c. Caustic soda and caustic potash
d. Sulphur dioxide and sulphur trioxide.

21. How many grams of H_3PO_4 is required to completely neutralize 120g of NaOH

a. 49
b. 98
c. 196
d. 9.8

22. The weight of oxygen required to completely react with 27 gms of 'Al' is

a. 8 gm
b. 16 gm
c. 32 m
d. 24 gm

23. If law of conservation of mass was to hold true, then 20.8 g of $BaCl_2$ on reaction with 9.8 g of H_2SO_4 will produce 7.3 g of HCl and $BaSO_4$ equal to:

a. 11.65 g
b. 23.3 g
c. 25.5 g
d. 30.6 g

24. One of the following combinations which illustrates the law of reciprocal proportions is:

a. N_2O_3, N_2O_4, N_2O_5
b. NaCl, NaBr, NaI
c. CS_2, CO_2, SO_2
d. PH_3, P_2O_3, P_2O_5

25. 23g of sodium will react with ethyl alcohol to give

a. 1 mole of H_2
b. 1/2 mole of H_2
c. 1 mole of O
d. 1 mole of NaOH

26. Hydrogen and oxygen combine to form H_2O_2 and H_2O containing 5.93% and 11.2% hydrogen respectively, the data illustrates:

a. Law of conservation of mass
b. Law of Constant proportions
c. Law of reciprocal proportions
d. Law of multiple proportions

27. Two elements X (of mass 16) and Y (of mass 14) combine to form compounds A, B and C. The ratio of different masses of Y which combine with a fixed mass of X in A, B and C is 1 : 3 : 5, if 32 parts by mass of X combines with 84 parts by mass of Y in B, then in C, 16 parts by mass of X will combine with;

a. 14 parts by mass of Y
b. 42 parts by mass of Y
c. 70 parts by mass of Y
d. 84 parts by mass of Y

Atomic & Molecular Masses

28. Insulin contains 3.4% sulphur by mass. What will be the minimum molecular weight of insulin?

a. 94.117 u
b. 1884 u
c. 941 u
d. 976 u

29. If we assume 1/24 th part of mass of carbon instead of 1/12 th part of it as 1 amu., mass of 1 mole of a substance will

a. Remain unchanged
b. get doubled
c. Get halved
d. can't be predicted

30. Boron has two isotopes 10B and 11B whose relative abundances are 20% and 80% respectively. Atomic weight of Boron is

a. 10
b. 11
c. 10.5
d. 10.8

31. 10 grams of each O_2, N_2 and Cl_2 are kept in three bottles. The correct order of arrangment of bottles containing decreasing number of Molecules.

a. O_2, N_2, Cl_2
b. Cl_2, N_2, O_2
c. Cl_2, O_2, N_2
d. N_2, O_2, Cl_2

32. Avogadro's number is the number of molecules present in

a. 1 g of molecule
b. 1atom of molecule
c. gram molecular mass
d. 1 litre of molecule

33. Maximum number of atoms are present in

a. 14 gms. of carbon monoxide
b. 2 gms. of hydrogen
c. 11.2 lit. of nitrogen at STP
d. 1.5 gm atoms of helium

34. One amu is equal to

a. 1.66×10^{-8} g
b. 1.66×10^{-4} g
c. 1.66×10^{-16} g
d. 1.66×10^{-24} g

35. The number of molecules present in one milli litre of a gas at STP is known as

a. Avogadro number

b. Boltzman number

c. Loschmidt number

d. Universal gas constant

36. Which of the following gases contain the same number of molecules as that of 16 grams of oxygen?

a. 16gm of O_3 b. 32 grams of SO_2

c. 16gm of SO_2 d. All

37. If the atomic mass unit 'u' were defined to be $\dfrac{1}{5}^{th}$ of the mass of an atom of C-12, what would be the atomic weight of nitrogen in amu or 'u' in this state? Atomic weight of N on conventional scale is 14:

a. 6.77 u   b. 5.834 u

c. 14 u   d. 23 u

38. A 100 g sample of Haemoglobin on analysis was found to contain 0.34% Fe by mass. If each haemoglobin molecule has four Fe^{2+} ions, the molecular mass of haemoglobin is- (Fe = 56 amu)

a. 77099.9 g   b. 12735 g

c. 65882 g   d. 96359.9 g

Mole Concept and Molar Masses

39. 1 g-atom of nitrogen represents:

a. 6.02×10^{23} N_2 molecules

b. 22.4 L of N_2 at S.T.P

c. 11.2 L of N_2 at S.T.P

d. 28 g of nitrogen

40. Which is correct for 10 g of $CaCO_3$?

a. It contains 1 g atom of carbon

b. It contains 0.3 g atoms of oxygen

c. It contains 12 g of calcium

d. It refers to 0.1 g equivalent of $CaCO_3$

41. The number of oxygen atoms present in 14.6 g of magnesium bicarbonate is:

a. $6 N_A$ b. $0.6 N_A$

c. N_A d. $\dfrac{N_A}{2}$

42. Which of the following has the highest mass?

a. 20 g of sulphur

b. 4 mol of carbon dioxide

c. 12×10^{24} atoms of hydrogen

d. 11.2 L of helium at N.T.P.

43. If isotopic distribution of C-12 ad C-14 is 98% and 2% respectively, then the number of C-14 atoms in 12 g of carbon is:

a. 1.032×10^{22} b. 3.01×10^{22}

c. 5.88×10^{23} d. 6.02×10^{23}

44. 5.6 L of a gas at S.T.P. weights equal to 8 g. The vapour density of gas is:

a. 32   b. 16

c. 8   d. 40

45. One atom of an element weighs 1.8×10^{-22} g, its atomic mass is:

a. 29.9 g   b. 18 g

c. 108.36 g   d. 154 g

46. If H_2SO_4 ionises as $H_2SO_4 + 2H_2O \rightarrow 2H_3O^+ + SO_4^{2-}$. Then total number of ions produced by 0.1 mol H_2SO_4 will be:

a. 9.03×10^{21} b. 3.01×10^{22}

c. 6.02×10^{22} d. 1.8×10^{23}

47. Which of the following will not have a mass of 10 g?

a. 0.1 mol $CaCO_3$

b. 1.51×10^{23} Ca^{2+} ions

c. 0.16 mol of CO_3^{2-} ions

d. 7.525×10^{22} Br atom

48. x L of N_2 at S.T.P. contains 3×10^{22} molecules. The number of molecules in x/2 L of ozone at S.T.P. will be:

a. 3×10^{22} b. 1.5×10^{22}

c. 1.5×10^{21} d. 1.5×10^{11}

49. A person adds 1.71 gram of sugar ($C_{12}H_{22}O_{11}$) in order to sweeten his tea. The number of carbon atoms added are: (mol mass of sugar = 342)

a. 3.6×10^{22} b. 7.2×10^{21}

c. 0.05   d. 6.6×10^{22}

50. The number of atoms present in 0.1 mole of P_4 (at. mass = 31) are:

a. 2.4×10^{24} atoms

b. Same as in 0.05 mol of S_8

c. 6.02×10^{22} atoms

d. Same as in 3.1g of phosphorus

51. Which one contains maximum number of molecules?

a. 2.5 g molecule of N_2

b. 4 g atom of nitrogen

c. 3.01×10^{24} atoms

d. 82 g of dinitrogen

52. Out of 1.0 g dioxygen, 1.0 g (atomic) oxygen and 1.0 g ozone, the maximum number of oxygen atoms are contained in:

a. 1.0 g of atomic oxygen

b. 1.0 g of ozone

c. 1.0 g of oxygen gas

d. All contain same number of atoms

53. The maximum volume at S.T.P. is occupied by:

a. 12.8 g of SO_2
b. 6.02×10^{22} molecules of CH_4
c. 0.5 mL of NO_2
d. 1g molecule of CO_2

54. If N_A is Avogadro's number, then the number of oxygen atoms in one g-equivalent of oxygen is:

a. N_A
b. $N_A/2$
c. $N_A/4$
d. $2N_A$

55. If 224 ml. of a triatomic gas has a mass of 1g at 273 K and 1 atm pressure, then the mass of one atom is:

a. 8.30×10^{-23} g
b. 6.24×10^{-23}
c. 2.08×10^{-23} g
d. 5.54×10^{-23} g

56. The rest mass of an electron is 9.11×10^{-31} kg. Molar mass of the electron is:

a. 1.5×10^{-31} kg mol^{-1}
b. 9.11×10^{-31} kg mol^{-1}
c. 5.5×10^{-7} kg mol^{-1}
d. 6.02×10^{23} kg mol^{-1}

57. A sample of ammonium phosphate, $(NH_4)_3 PO_4$, contains 3.18 moles of hydrogen atoms. The number of moles of oxygen atoms in the sample is:

a. 0.265
b. 0.795
c. 1.06
d. 3.18

58. What is the total number of atoms present in 25.0 mg of camphor, $C_{10}H_{16}O$?

a. 9.89×10^{19}
b. 6.02×10^{20}
c. 9.89×10^{20}
d. 2.67×10^{21}

59. Which of the following samples contains the largest number of atoms?

a. 1 g of CO_2
b. 1 g of C_8H_{18}
c. 1 g of C_2H_6 (g)
d. 1 g of LiF (s)

60. 4.0 g of caustic soda (NaOH) (mol mass 40) contains same number of sodium ions as are present in-

a. 10.6 g of Na_2CO_3 (mol. mass 106)
b. 58.5 g of NaCl (Formula mass 58.5)
c. 100 ml of 0.5 M Na_2SO_4 (Formula mass 142)
d. 1mol of $NaNO_3$ (mol. mass 85)

61. Total number of atoms present in 64 gm of SO_2 is -

a. $2 \times 6.02 \times 10^{23}$
b. 6.02×10^{23}
c. $4 \times 6.02 \times 10^{23}$
d. $3 \times 6.02 \times 10^{23}$

62. The total number of protons, electrons and neutrons in 12gm of $_6C^{12}$ is -

a. 1.084×10^{25}
b. 6.022×10^{23}
c. 6.022×10^{22}
d. 18

63. Number of Ca^{+2} and Cl^- ion in 111 g of anhydrous $CaCl_2$ are -

a. $N_A, 2N_A$
b. $2N_A, N_A$
c. N_A, N_A
d. None

64. The maximum volume at N.T.P. is occupied by-

a. 12.8 gm of SO_2
b. 6.02×10^{22} molecules of CH_4
c. 0.5 mol of NO_2
d. 1 gm-molecule of CO_2

65. Number of moles of water in 488 g of $BaCl_2.2H_2O$ are - (Ba = 137)

a. 2 moles
b. 4 moles
c. 3 moles
d. 5 moles

66. 4.4 g of CO_2 and 2.24 litre of H_2 at STP are mixed in a container. The total number of molecules present in the container will be -

a. 6.022×10^{23}
b. 1.2044×10^{23}
c. 2 moles
d. 6.023×10^{24}

67. One mole of nitrogen gas has volume equal to

a. 1 litre of nitrogen at S.T.P.
b. 32 litre of nitrogen at S.T.P.
c. 22.4 litre of nitrogen at S.T.P.
d. 11.2 litre of nitrogen at S.T.P.

68. The number of atoms present in 0.1 mole of P_4.

a. 2.4×10^{23} atom (approx)
b. Same as in 0.05 mole of S_8
c. Same as in 12.4 g of P_4
d. 2.4×10^{24} atom (approx)

Percentage Composition

69. The percentage of C, H and N in an organic compound are 40 %, 13.3% and 46.7% respectively then empirical formula is

a. $C_3H_{13}N_3$
b. CH_2N
c. CH_4N
d. CH_6N

70. B_1 g of an element gives B_2 g of its chloride, the equivalent mass of the element is:

a. $\dfrac{B_1}{B_2 - B_1} \times 35.5$
b. $\dfrac{B_2}{B_2 - B_1} \times 35.5$
c. $\dfrac{B_2 - B_1}{B_1} \times 35.5$
d. $\dfrac{B_2 - B_1}{B_2} \times 35.5$

71. 60 g of a compound on analysis gave 24 g C, 4 g H and 32 g O. The empirical formula of the compound is:

a. $C_2H_4O_2$
b. $C_2H_2O_2$
c. CH_2O_2
d. CH_2O

72. A compound made of two elements A and B are found to contain 25% A (at mass 12.5) and 75% B (at mass 37.5). The simplest formula of the compound is:

a. AB
b. AB_2
c. AB_3
d. A_3B

73. Insulin contains 3.4% sulphur. What will be the minimum molecular weight of insulin?
   a. 94.117
   b. 1884
   c. 941.176
   d. 976

74. 400 mg of capsule contains 100 mg of ferrous fumarate. The percentage of Fe present in the capsule is approximately: (formula of ferrous fumarate is $(CHCOO)_2 Fe$).
   a. 8.2%
   b. 25%
   c. 16%
   d. Unpredictable

75. Simplest formula of compound containing 50% of element X (at mass 10) and 50% of element Y (at mass 20) is:
   a. XY
   b. X_2Y
   c. XY_2
   d. X_2Y_3

76. A compound having the empirical formula (C_3H_4O) has a molecular mass of 170 ± 5. The molecular formula of it compound is:
   a. C_3H_4O
   b. $C_6H_8O_2$
   c. $C_6H_{12}O_3$
   d. $C_9H_{12}O_3$

77. Two oxides of a metal contains 50% and 40% metal (M) respectively. If the formula of first oxide is MO_2, the formula of second oxide will be:
   a. MO_2
   b. MO_3
   c. M_2O
   d. M_2O_5

78. The vapour density of gas A is four times that of B. If molecular mass of B is M, then molecular mass of A is:
   a. M
   b. 4 M
   c. $\dfrac{M}{4}$
   d. 2 M

79. A metal nitride M_3N_2 contains 28% of nitrogen. The atomic mass of metal M is:
   a. 24
   b. 54
   c. 9
   d. 87.62

80. A container of volume V, contains 0.28 g of N_2 gas. If same volume of an unknown gas under similar conditions of temperature and pressure weights 0.44 g, the molecular mass of gas is:
   a. 22
   b. 44
   c. 66
   d. 88

81. A gaseous hydrocarbon on complete combustion gives 3.38 g of CO_2 and 0.690 g of H_2O and no other products. The empirical formula of hydrocarbon is:
   a. CH
   b. CH_2
   c. CH_3
   d. The data is not complete

82. The percentage of Carbon in CO_2 is
   a. 27.27%
   b. 29.27%
   c. 30.27%
   d. 26.97%

83. The haemoglobin from red blood corpuscles of most mammals contain approximately 0.33% of iron by mass. The molecular mass of haemoglobin is 67200. The number of iron atoms in each molecule of haemoglobin is:
   a. 3
   b. 4
   c. 2
   d. 6

84. On analysis a certain compound was found to contain iodine and oxygen in the ratio of 254 g of iodine (at mass 127) and 80 g oxygen (at mass 16). What is the formula of compound?
   a. IO
   b. I_2O
   c. I_5O_3
   d. I_2O_5

85. 0.5 mol of potassium ferrocyanide contains carbon equal to: (Formula of potassium ferrocyanide is $K_4[Fe(CN)_6]$).
   a. 1.5 mol
   b. 36 g
   c. 18 g
   d. 3.6 g

86. 14 g of element X combine with 16g of oxygen. On the basis of this information, which of the following is a correct statement :
   a. The element X could have an atomic weight of 7 and its oxide formula XO
   b. The element X could have an atomic weight of 14 and its oxide formula X_2O
   c. The element X could have an atomic weight of 7 and its oxide is X_2O
   d. The element X could have an atomic weight of 14 and its oxide is XO_2

87. A compound has 20% of nitrogen by weight. If one molecule of the compound contains two nitrogen atoms, the molecular weight of the compound is
   a. 35
   b. 70
   c. 140
   d. 280

Stoichiometry & Stoichiometric Calculation

88. 'X' litres of carbon monoxide is present at STP. It is completely oxidized to CO_2. The volume of CO_2 formed is 11.207 litres at STP. What is the value of 'X' in litres?
   a. 22.414
   b. 11.207
   c. 5.6035
   d. 44.828

89. The volume of phosgene formed at STP when 11.2 lit of chlorine reacts with carbon monoxide is
   a. 11.2 lit
   b. 22.4 lit
   c. 5.6 lit
   d. 44.8 lit

90. The moles of O_2 required for reacting with 6.8 g ammonia. $(.....NH_3+....O_2\rightarrow.....NO+.....H_2O)$ is:
   a. 5
   b. 2.5
   c. 1
   d. 0.5

91. What mass of $CaCl_2$ in grams would be enough to produce 14.35 gm of AgCl?
   a. 5.55 g
   b. 8.29 g
   c. 16.59 g
   d. 10 g

92. What weight of sodium hydroxide is required to neutralise 100 ml of 0.1N HCl ?

a. 4 g   b. 0.4 g

c. 0.04 g   d. 40 g

93. H_2O_2 is sold as a solution of approximately 5.0 g H_2O_2 per 100 ml of the solution. The molarity of this solution is approximately

a. 0.15 M   b. 1.5 M

c. 3.0 M   d. 3.4 M

94. The amount of oxalic acid (eq.wt.63) required to prepare 500 ml of its 0.10 N solution is

a. 0.315 g   b. 3.150 g

c. 6.300 g   d. 63.00 g

95. The molarity of pure water is

a. 100 M   b. 55.6M

c. 50 M   d. 18 M

96. The mass of 70% H_2SO_4 by mass is required for neutralisation of 1 mole of NaOH is:

a. 65   b. 98

c. 70   d. 54

97. If potassium chlorate is 80% pure then 48 g of oxygen would be produced from:

a. 153.12 g of $KClO_3$ b. 120 g of $KClO_3$

c. 20 g of $KClO_3$ d. 90 g of $KClO_3$

98. Density of a solution containing x% by mass of H_2SO_4 is y. The normality is

a. $\dfrac{xy \times 10}{98}$ b. $\dfrac{xy \times 10}{98y} \times 2$

c. $\dfrac{xy \times 10}{98} \times 2$ d. $\dfrac{x \times 10}{98y}$

99. Mass percentage (w/w) of ethylene glycol ($HOCH_2 - CH_2OH$) in a aqueous solution is 20, then mole fraction of solute is

a. 0.5   b. 0.067

c. 0.1   d. 0.4

100. Number of gram equivalents of solute in 100 ml of 5 N HCl solution is

a. 50   b. 500

c. 5   d. 0.5

101. If 1.26 grams of oxalic acid is dissolved in 250 ml of solution then its normality is

a. 0.05   b. 0.04

c. 0.02   d. 0.08

102. 100ml of ethylalcohol is made upto a litre with distilled water. If the density of C_2H_5OH is 0.46 gm/ml. Then its molality is

a. 0.55 m   b. 1.11m

c. 2.22 m   d. 3.33m

103. A solution of 0.1 mole of a metal chloride MCl_x required 500 mL of 0.6 molal $AgNO_3$ solution for complete ppt. The value of x is:

a. 5   b. 4

c. 3   d. 1

104. If 20 g of $CaCO_3$ is treated with 100 ml 20% HCl solution. The amount of CO_2 produced is:

a. 22.4l g   b. 8.8 g

c. 2.2 g   d. 8 l

105. The mass of $CaCO_3$ required to react with 25 mL of 0.75 molar HCl is:

a. 0.94 g   b. 0.68 g

c. 0.76 g   d. 0.52 g

106. 2 moles of H_2S and 11.2 L of SO_2 at N.T.P. reacts to form x moles of sulphur. The value of x is:

a. 1.5   b. 3.5

c. 7.8   d. 12.7

107. Sulphuryl chloride (SO_2Cl_2) reacts with H_2O to give a mixture of H_2SO_4 & HCl. Aqueous solution of 1 mole SO_2Cl_2 will be neutralised by:

a. 3 moles of NaOH

b. 2 moles of $Ca(OH)_2$

c. Both (a) & (b)

d. None of these

108. A sample of pure compound contains 1.15 g of sodium, 3.01×10^{22} atoms of carbon and 0.1 mol of oxygen atom. Its empirical formula is:

a. Na_2CO_3 b. $NaCO_2$

c. $Na_2 CO$ d. Na_2CO_2

109. If 0.30 mol of zinc are added to 0.52 mol of HCl, the moles of H_2 formed are:

a. 0.52   b. 0.30

c. 0.26   d. 0.60

110. The specific gravity of 98% H_2SO_4 is 1.8 g/cc. 50 ml of this solution is mixed with 1750 ml of pure water. Molarity of resulting solution is

a. 0.2 M   b. 0.5 M

c. 0.1 M   d. 1 M

ABOUT PHYSICS WALLAH

Alakh Pandey is one of the most renowned faculty in NEET & JEE domain's Physics. On his YouTube channel, Physics Wallah, he teaches the Science courses of 11th and 12th standard to the students aiming to appear for the engineering and medical entrance exams.

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