**Theory:**

The schematic of a reactor is shown in the figure below:

NC (Inlet flow) NC (Outlet flow)

Reactor with specific T, P

The mass balance for the reactor is written as follows:

The first thing we must know is how many reactions we need to determine mass balance?

Finding the number of a reaction:

And [A] term is the rank of A matrix (matrix of species). It shows how many independent variables we have.

The next issue is how to find . With Gibbs-free energy versus … plot, we can determine with calculate global minimum of this plot, which mathematically is described as follows:

Next step is to calculate partial derivative of this equation:

Respect to Lewis’s equation we have:

The ratio “” described as activity coefficient (ai­)

Thus, the equation … will be converted to:

Additionally, the equation … will be converted to:

By knowing the fugacity of each component, we can determine the fugacity of mixture. This is very complicated from the numerical point of view.

1. **T effect @ constant P:**

Vant Hoff equation:

If , it indicates that we have an exothermic reaction. The increasing of T leads to decreasing .

Otherwise, we have an endothermic reaction. The decreasing of T leads to increasing .

1. **P effect @ constant T**

Suppose a reaction like this:

The equilibrium constant will be expressed as follows:

if c+d > a+b: With increasing pressure, the term will be decreased.

Elseif c+d < a+b: With increasing pressure, the term will be increased.

Else: With increasing pressure, the term will be constant.

The schematic of all system is shown in the figure below:

A diagram of a flowchart

Description automatically generated