

Source: [KBiologyMasterIndex](#)

1 | Alright, let's talk about water.

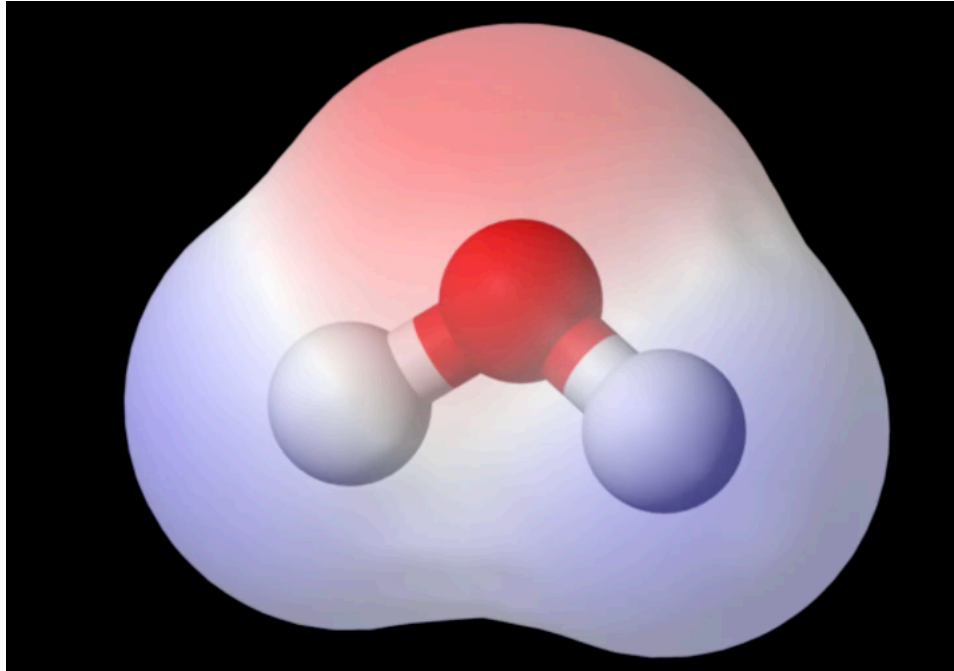


Figure 1: Watah!

Intra-Molecular water bonds

- As we know (or looked up from the PTable)
- Hydrogen has electronegativity of 2.20, and oxygen has EN 3.44
- The difference $>0.4 <1.7$ makes these bonds **polar covalent**

See [KBhBIO101BondingReview](#), bonding review

Why is ice less dense?

- Freezing usually bring new ordering of molecules to make it paced
- But! Not the case for normal water because of hydrogen bond
 - Strong 6-ring Hydrogen bonds in ice prevents shrinking
 - Empty space filled with air, making it less dense

Why do I care?

Ice floats! Causing beautiful ice effects that the ecosystem uses

- Pseudo-landmass
 - Antarctica

- Animal habitat
- Ice reflect incoming light (“albedo”)
 - Bounces energy off
 - Cools the water

Inter-Molecular water bonds properties

See [|KBhBIO101PropsOfWater|](#), properties of water.

2 | So, why is water the chosen liquid?

See also [|KBhBIO101PropsOfWater|](#) Properties of Water

- Liquid at Earth temperatures
- Sticky => strong bonds that help water hold together + resist change in temperature (hence why AlcoholLand cannot exist).
 - Varing te (incomplete #todo-houjun @jemoka)
- Is universal solvent