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1 | Bio-Molecules Quiz Review

Bonding in organic compounds, a review.

common nonpolar bonds

Carbon-carbon
Carbon-hydrogen
Carbon-sulfur

Common dipole interactions

Carbon-nitrogen $\delta^+ - \delta^-$ Carbon-oxygen $\delta^+ - \delta^-$
Nitrogen-oxygen $\delta^+ - \delta^-$ Hydrogen-oxygen $\delta^+ - \delta^-$

Common ionic interactions

they come from acid-base interactions.

However, sometimes they are permanent. Look at the amino acid chart for those.

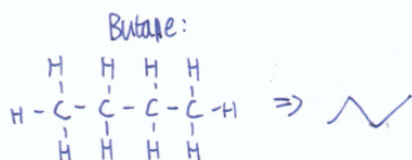
why hydrogen bonding is excellent

hydrogen bonding allows stronger dipole-dipole bonds than dipole-dipole bonds. They are still good ol covalent bonds.

These bonds basically combines Hydrogen w/ the most electronegative atoms.

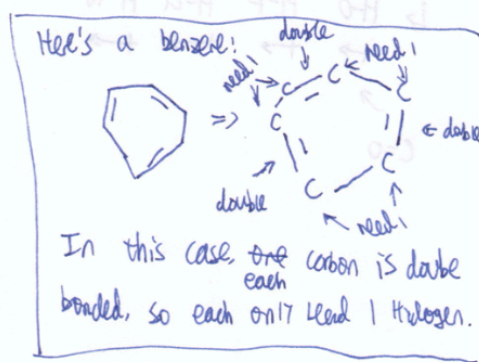


Reading a line-angle representation.



In this type of representations, start with a line. Bend the line at every carbon.

Now, it is assumed that carbon is not going to just be happy with $\text{C}-\text{C}-\text{C}-\text{C}$.



so, we still the missing orbitals with hydrogen.

$\text{C}-\text{C}-\text{C}-\text{C}$ ← need 3

↑ need 3 ↑ need 2

Figure 1: Screen Shot 2020-10-09 at 11.58.55 AM.png