

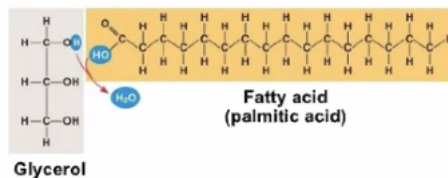
Source: [KBhBIO101Macromolecules](#)

#disorganized

Firstly, LDF.

Pentaine => Carbon-Carbon bonds has 0 EN difference w/ even distribution of electrons

- Electrons are randomly within their 3D shape
- So, when electrons are accidentally concentrated, a dipole moment will temporarily form ***
- Pentaine!
 - Non polar
 - Only have LDF
 - So, water's dipole moment isn't very attractive anymore
 - Larger surface area of pentain makes it happier bonding with each other instead of water (more LDF opportunities with bigger surface area of pentain)
- Definition: lipids — dominantly hydrophobic biological molecules ("fats")



Single fatty acids

- Fatty acids — Glycerol
- Fat (a.k.a. adipose tissue) — Triglycerol (3) — 3 gl
- Saturated vs. unsaturated fats
 - No Double bonds in left carbon chain: saturated fats, (butter)
 - Double bonds in left carbon chain: unsaturated fats, (olive oil)
- A question: given that saturated fats have no double bonds, why does it melt later than the unsaturated fats?
 - Double bonds, due to their caused VESPR geometry (and hence the -1 hydrogen), are curved. This makes it harder to stack together, causing a lower melting point
 - Single bonds, due to their caused VESPR geometry, is flat. This makes them easier to stack together, causing a higher melting point.

Phosolipids = 2 fatty acids + phosphate group

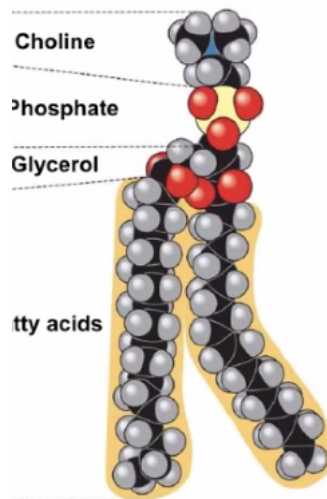
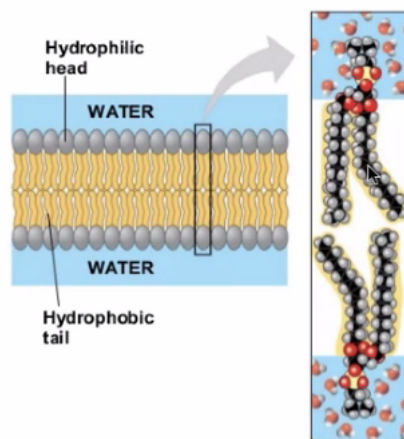


Figure 1: Screen Shot 2020-09-09 at 3.15.41 PM.png



Many of these will end up with membrane

- Head hydrophobic
- Tail hydrophilic
- Causes a string

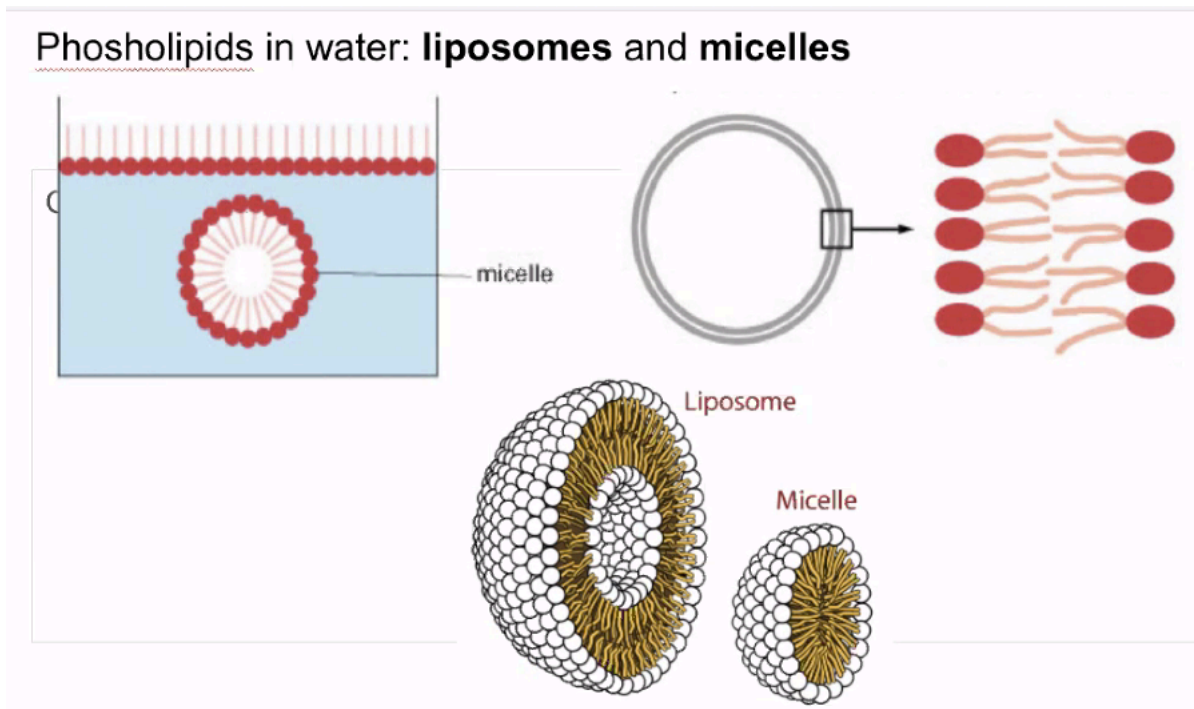


Figure 2: Screen Shot 2020-09-09 at 3.11.54 PM.png