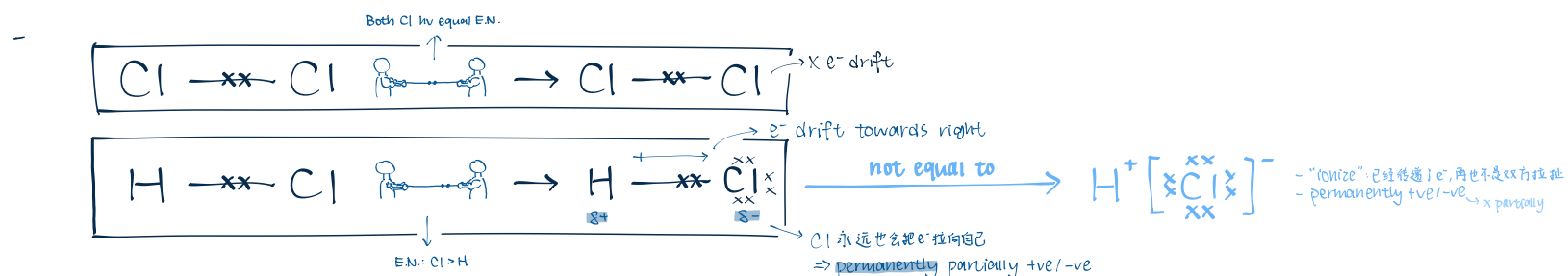


E.N. and polar bonds

1 Electronegativity

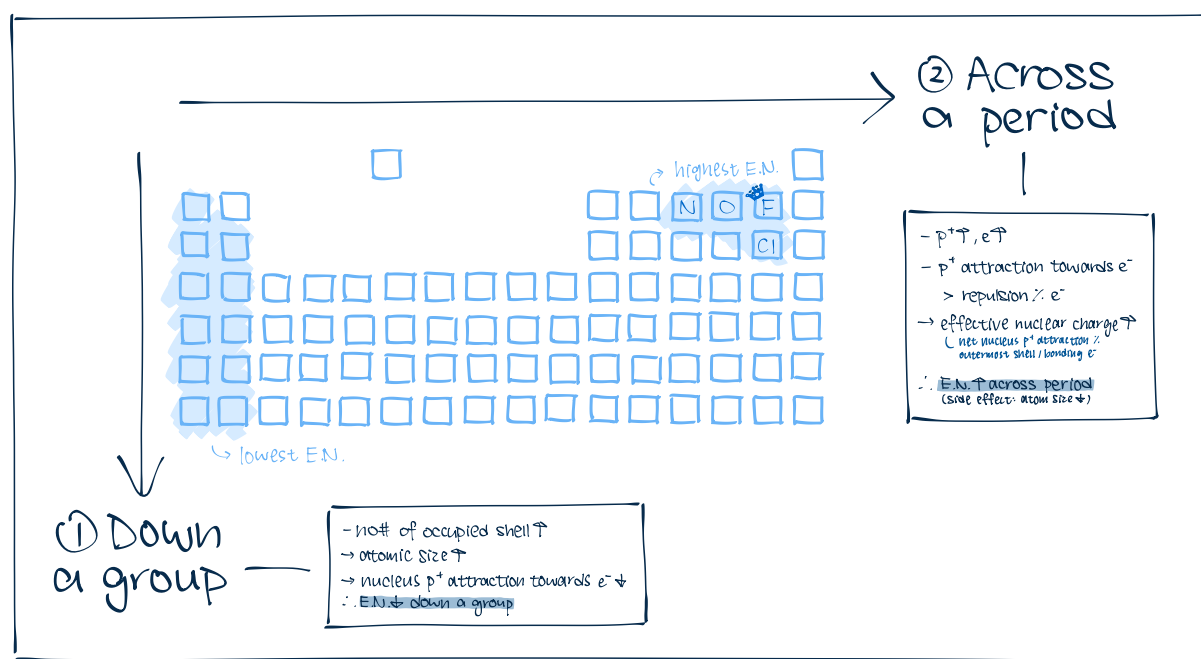
- ability of an atom to withdraw bonding e^- towards itself in a covalent bond



- factors affecting E.N.

- E.N. = ability of atom to attract bonding $e^- \rightarrow$ atom 里什么真正吸引 e⁻? : proton

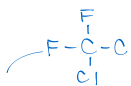
\rightarrow E.N. = attraction strength of nucleus p^+ towards bonding e^- (diagram of nucleus and electron)



2 Polar bonds

WHAT ARE THEY?

- uneven distribution of bonding e^- in covalent bond
- diff. atoms in a bond \rightarrow diff. E.N. \rightarrow polar
- \uparrow E.N. diff. of 2 atoms \rightarrow \uparrow bond polarity
- eg. identify polar bonds + compare polarity of them for CF_2Cl_2
 - > C-F bond & C-Cl bond
 - C & F atoms, C & Cl atoms have different electronegativities
 - > polarity: C-F bond > C-Cl bond
 - diff. in E.N.: C, F > C, Cl — E.N.: F > Cl > C (highest in all elements)



REPRESENTATION OF POLAR BONDS

