Lonic equations

I What are ionic equations?

- chemical equ. which only include ions formed / changed
- 一 ie 只包括真正参与reaction 的物质
- Steds
 - 1. 写一次已经 balanced 的 equation
 - 2. 把 angueous 的物质折开成 ion
 - 3. 把沒改变过的 ion 删走 > spectator ion
 - 4. 確保內边 charge 的数量相等

2 Examples

Mg + Al2(SO4)3.

- - $RHS = 3 \times (+2) + 0 = +6$

Ca+HC1-

- Ca(s) + 2HC/caq) -> CaC/2 (aq) + H2 cg)
- ⇒ Ca+2H⁺ → Ca²⁺ + H₂ → Cl⁻ 沒要过, H₂不能抓
- LHS = 2x(+1) = +2RHS = 1 x (+2) = +2

Pb(NO3)2+Na2SO4-

- Pb(NO3)2 (ag) + NO12 SO4 (ag) -> PbSO4 (s) + 2NO1 NO3 (ag)
- ⇒ $Pb^{2+} + SO_4^{2-}$ → $PbSO_4$ → $PbSO_4$ $PbSO_4$
- LHS = (+2) + (-2) = 0RHS=0