

Describing O.C.s

1 Colour / smell of gases

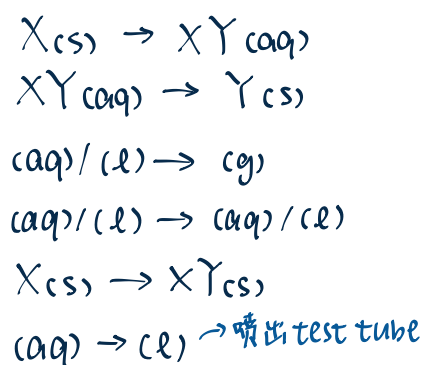
	Cl_2	very pale yellowish green gas	} Toxic. - If the exp. produces these gases, it should be performed in a fume cupboard.
	Br_2	brown gas	
	I_2	purple gas	
requires heat ($\because \text{Br}_2$ is liquid in room conditions)	NO_2	brown gas	
	SO_2	choking / pungent smell	
	NH_3	choking / pungent smell	

2 Colour of salts (s)



- both colourless in aq → usually white
- either one coloured → follow its colour
- both coloured → mix colours / choose more distinct colour

2 Observable changes



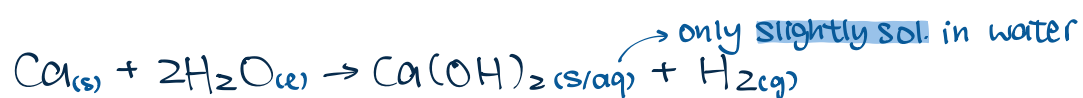
X dissolves
 <colour> solid deposits
 <colour> bubbles evolve
 solⁿ turns from <colour> to <colour>
 <colour> solid turns <colour>
 <colour> fumes evolve

3 Examples



- $\text{Cu}_{(s)} \rightarrow \text{Cu}^{2+}_{(aq)}$
- $\rightarrow \text{Cu}^{2+}_{(aq)}$
- $\rightarrow \text{SO}_{2(g)}$
- $\rightarrow \text{SO}_{2(g)}$

copper dissolves → 只能说有什么 reactant, product 要用 colour + state 形容
 solution turns from colourless to blue
 colourless gas bubbles evolve
 choking smell



- $\text{Ca}_{(s)} \rightarrow \text{Ca}^{2+}_{(aq)}$
- $\text{Ca}_{(s)} \rightarrow \text{Ca}(\text{OH})_{2(aq)}$
- $\rightarrow \text{H}_{2(g)}$

calcium dissolves
 white solid deposits
 colourless gas bubbles evolve