

Describing O.C.s

1 Colour / smell of gases

Cl_2	very pale yellowish green gas
Br_2	brown gas
NO_2	brown gas
I_2	purple gas
SO_2	choking / pungent smell
NH_3	choking / pungent smell

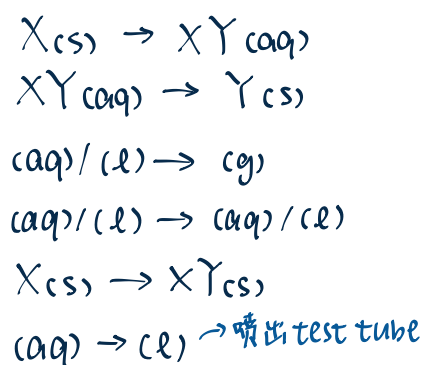
requires heat
(Br_2 is gas in room conditions)

2 Colour of salts (s)



- both colourless in aq → usually white
- either one coloured → follow its colour
- both coloured → mix colours / choose more distinct colour

2 Observable changes

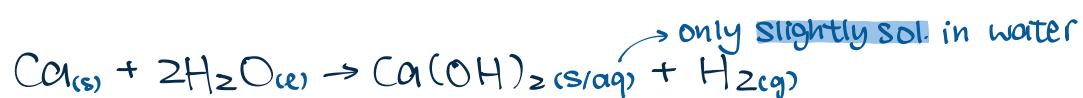


X dissolves
<colour> solid deposits
<colour> bubbles evolve
solⁿ turns from <colour> to <colour>
<colour> solid turns <colour>
<colour> fumes evolve

3 Examples



- $\text{Cu}_{(s)} \rightarrow \text{Cu}^{2+}_{(aq)}$ copper dissolves → 只能说有什么 reactant, product 要用 colour + state 形容
- $\rightarrow \text{Cu}^{2+}_{(aq)}$ solution turns from colourless to blue
- $\rightarrow \text{SO}_{2(g)}$ colourless gas bubbles evolve
- $\rightarrow \text{SO}_{2(g)}$ choking smell



- $\text{Ca}_{(s)} \rightarrow \text{Ca}^{2+}_{(aq)}$ calcium dissolves
- $\text{Ca}_{(s)} \rightarrow \text{Ca}(\text{OH})_{2(aq)}$ white solid deposits
- $\rightarrow \text{H}_{2(g)}$ colourless gas bubbles evolve