

pH and molarity

1 Formulas

$$pH = -\log [H^+] \quad \text{molarity}$$

PH为数值, 没单位

$$molarity = \frac{\text{mole}}{\text{volume}} \quad \text{mol} \quad dm^3$$

1 dm³ = 1000 cm³

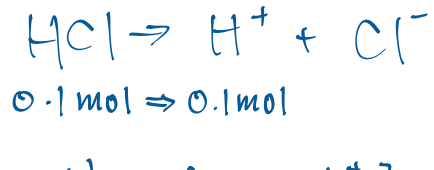
2 Mole concept 金句

- 加水不影响 no. of mole
- 加水至 n 倍 volume → molarity 变至 1/n 倍 (no. of mole 不变)
- 抽去倍 volume 的溶液 → no. of mole 变至 1/n 倍 (molarity 不变)

3 Questions

LEVEL I

0.1 mole of HCl completely dissolves in 200 cm³ distilled water. Find the pH of solution.

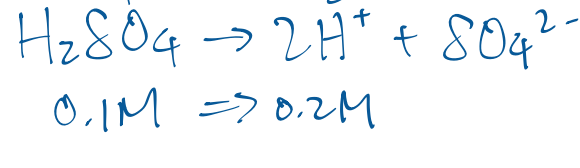


$$0.1 \text{ mol} \Rightarrow 0.1 \text{ mol}$$

$$pH = -\log [H^+] = -\log (0.2) = 0.301$$

LEVEL II

50 cm³ of 0.1M H₂SO₄ → find the solution pH.



$$0.1M \Rightarrow 0.2M$$

$$pH = -\log [H^+] = -\log 0.2 = 0.699$$

LEVEL III

30 cm³ of 0.1M H₂SO₄ vs 50 cm³ of 0.1M H₂SO₄

Which one is more acidic?

same

$$\text{Both } [H^+] = 0.1M \times 2 = 0.2M$$

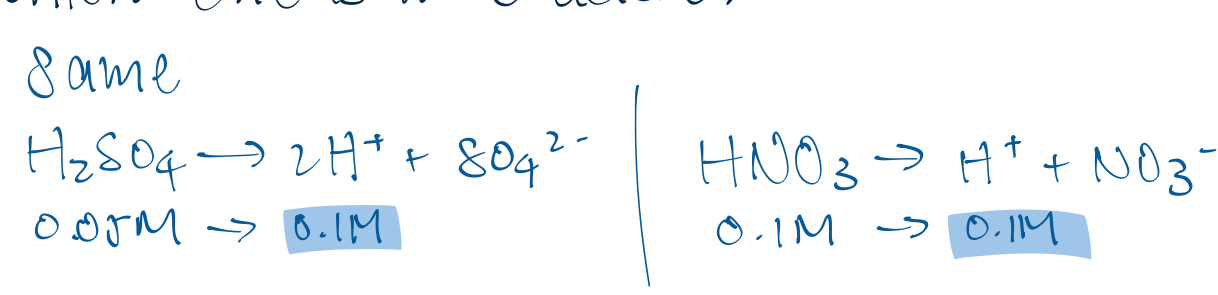
molarity 不变 → volume 增加 H⁺ 也会增加.

总体浓度还是一样

30 cm³ of 0.05M H₂SO₄ vs 80 cm³ of 0.1M HNO₃

Which one is more acidic?

same



$$0.05M \Rightarrow 0.1M \quad 0.1M \Rightarrow 0.1M$$

LEVEL IV

80 cm³ of 0.1M HCl is mixed w/ 40 cm³ of 0.1M HCl

Find the resulting pH.

molarity 一样, volume 不同

混合在一起 molarity 不变

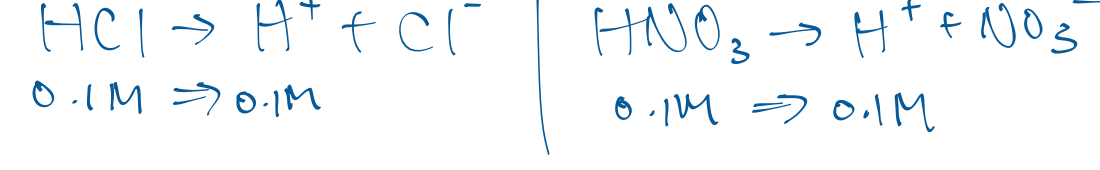


$$0.1M \Rightarrow 0.1M$$

$$pH = -\log [H^+] = -\log 0.1 = 1$$

80 cm³ of 0.1M HCl is mixed with 50 cm³ of 0.1M HNO₃

先算两边 H⁺ 的浓度

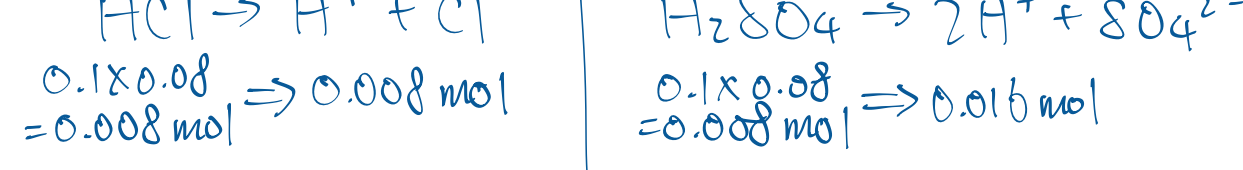


$$0.1M \Rightarrow 0.1M \quad 0.1M \Rightarrow 0.1M$$

$$pH = -\log [H^+] = -\log 0.1 = 1$$

80 cm³ of 0.1M HCl is mixed w/ 80 cm³ of 0.1M H₂SO₄

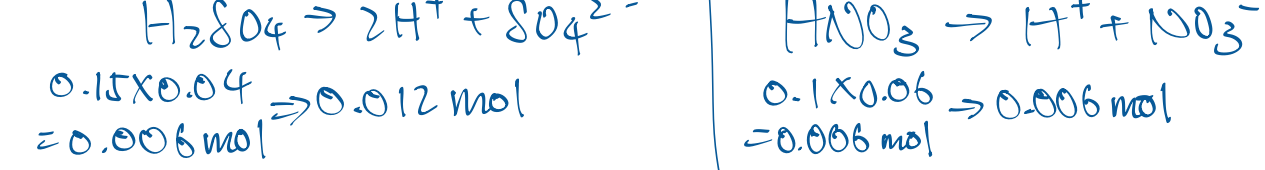
这道题得先算 H⁺ 的 no. of mole (∵ 两种 acid 的 [H⁺] 不同)



$$0.1 \times 0.08 \Rightarrow 0.008 \text{ mol} \quad 0.1 \times 0.08 \Rightarrow 0.016 \text{ mol}$$

$$pH = -\log [H^+] = -\log \frac{0.008 + 0.016}{0.016} = 0.824$$

40 cm³ of 0.15M H₂SO₄ is mixed w/ 60 cm³ of 0.1M HNO₃



$$0.15 \times 0.04 \Rightarrow 0.012 \text{ mol} \quad 0.1 \times 0.06 \Rightarrow 0.006 \text{ mol}$$

$$pH = -\log [H^+] = -\log \frac{0.012 + 0.006}{0.1} = 0.745$$

LEVEL V

The pH of 50 cm³ HCl is 1.

How much water do we need to add to increase its pH value to 3?

$$\begin{aligned} \text{Original molarity} \\ 1 = -\log [H^+] \\ [H^+] = 0.1M \end{aligned}$$

$$\begin{aligned} \text{required molarity} \\ 3 = -\log [H^+] \\ [H^+] = 0.001M \end{aligned}$$

$$0.1M \times 0.05 \text{ dm}^3 = 0.005 \text{ mol}$$

$$\frac{0.005 \text{ mol}}{0.05 \text{ dm}^3 + v} = 0.001M$$

$$0.005 = 0.0005 + 0.001v$$

$$v = 4.95 \text{ dm}^3$$

The pH of 10 cm³ H₂SO₄ is 2.

How much water do we need to add to increase its pH value to 4?

$$\begin{aligned} \text{Original molarity} \\ 2 = -\log [H^+] \\ [H^+] = 0.01M \end{aligned}$$

$$\begin{aligned} \text{required molarity} \\ 4 = -\log [H^+] \\ [H^+] = 0.0001M \end{aligned}$$

$$0.01M \times 0.01 \text{ dm}^3 = 0.0001 \text{ mol} \rightarrow \text{不用因为 } H_2SO_4 \text{ 而来?}$$

(已经计算了 H⁺ 的浓度, 而非 H₂SO₄ 的浓度)

$$\frac{0.0001 \text{ mol}}{0.01 \text{ dm}^3 + v} = 0.0001M$$

$$1 = 0.01 + v \rightarrow \text{约 } 0.001$$

$$v = 0.99 \text{ dm}^3$$

LEVEL VI

concentration 有两种单位

mol dm⁻³ (molarity)

g cm⁻³ (density)

⇒ 要懂得两种单位之间的互換

0.1M HCl → Find its density in g cm⁻³.

$$0.1M = \frac{0.1 \text{ mol}}{1 \text{ dm}^3} = \frac{0.1(35.5+1)g}{1000 \text{ cm}^3}$$

$$= 0.00365 \text{ g cm}^{-3}$$

80 cm³ of 0.1M impure HCl contains 83% HCl by mass.

Find the density of HCl.

$$\text{mole of acids} = 0.08 \times 0.1 = 0.008 \text{ mol}$$

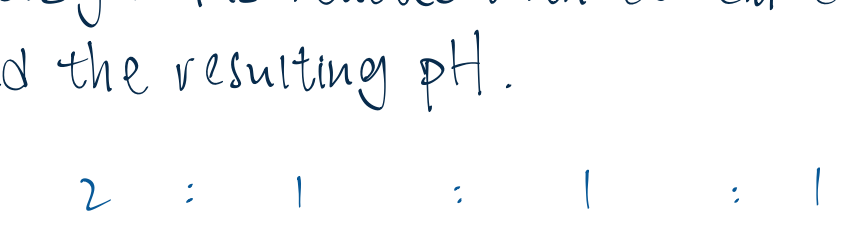
$$\text{mass of HCl} = \frac{0.08 \times 83\% \times 36.5}{\text{mole} \quad \text{Mr}} = 0.242 \text{ g}$$

$$\text{density of HCl} = \frac{0.242}{80} = 0.00303 \text{ g cm}^{-3}$$

LEVEL VII

If 0.3g Na is reacted with 600 cm³ of 0.1M H₂SO₄.

find the resulting pH.



$$\frac{0.3g}{23} \quad 0.6 \times 0.1$$

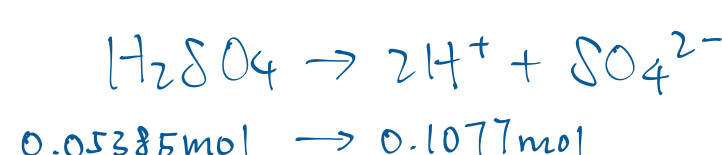
$$= 0.0123 \text{ mol} \quad = 0.06 \text{ mol}$$

$$\text{required } H_2SO_4 = 0.00615 \text{ mol}$$

H₂SO₄ is excess,

$$\text{unused } H_2SO_4 = 0.06 - 0.00615$$

$$= 0.05385 \text{ mol}$$

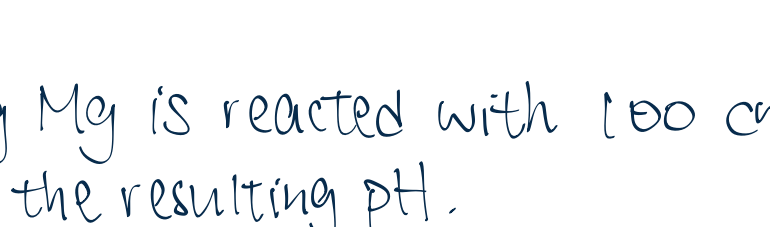


$$0.05385 \text{ mol} \rightarrow 0.1077 \text{ mol}$$

$$pH = -\log [H^+] = -\log \frac{0.1077}{0.6} = 0.746$$

If 3g Mg is reacted with 100 cm³ of 0.1M HCl,

find the resulting pH.



$$\frac{3g}{24.3} \quad 0.1 \times 0.1$$

$$= 0.0123 \text{ mol} \quad = 0.01 \text{ mol}$$

$$\text{required } H_2SO_4 = 0.0246 \text{ mol}$$

All H₂SO₄ is used up

pH = 7