

Mixed question types

1 Question 1

- Explain why K reacts w/ dilute acids.
- \therefore K reacts w/ water.
Range of metal reacting w/ acid > water
 \therefore K reacts w/ dilute acids

2 Question 2

- Explain which is more reactive, K or Ag.
- K
To produce metal oxide, K doesn't require heat while Ag needs strong heating.

3 Question 3

- According to the following table, arrange reactivities of W, X, Y, Z ascendingly.

	W	X	Y	Z
+CuSO ₄ aq	colourless gas bubbles	reddish brown solid deposits	no O.C.	no O.C.
heating	W ₂ O no O.C.	XO no O.C.	Y ₂ O silvery solid deposits	ZO no O.C.

- Y, Z, X, W
- (X, Y, Z) < W
W reacts w/ water, others do not
- (Y, Z) < X
X is a stronger reducing agent than Cu, others are not
- Y < Z
Y₂O decomposes and gives Y upon heating, ZO does not