Atoms

1 What are atoms?

- small particles making up everything
- very tiny & lightweight
- different atoms have different sizes / weights

2 Structure of atoms

PARTICLES

(atoms) same no. → overall neutral

	name	symbol	relative mass	charge
-	Proton	p	1	+1
-	Neutron	n	1	0
-	Electron	e^-	0 (negligible)	-1

STRUCTURE

A potassium atom (19 electrons)

Nucleus

- contains protons & neutrons
  - ↳ positively charged
- tiny & extremely dense (diameter = $\frac{1}{50000}$ of atom diameter)

Electron shells

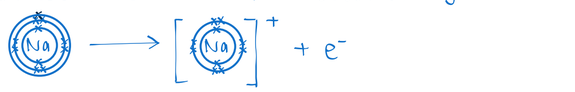
- certain fixed orbits surrounding nucleus which electrons move in
- max number of e^- a shell holds = $2n^2$ → 2 ⇒ 8 ⇒ 18 ⇒ 32

☆ 壳不一定要被填满

Electronic arrangement

- Represented by numbers
  - → eg. $_{13}Al$ → 2, 8, 3 (1st shell, 2nd shell, 3rd shell)
- Represented by electron diagrams
  - → eg. oxygen atom
  - → eg2. oxide ion

- octet rule
  - → All atoms tend to attain stable electronic arrangement (duplet / octet) by gaining / losing e^- → ions
  - → eg. Na loses 1 outermost shell e^- → octet arrangement → Na^+ ion



3 Representing atoms

FULL ATOMIC SYMBOL