Redissolving of metal hydroxides

I what is redissolve?

- precipitate dissolve when excess reactant is used : further reactions

	All # level + NaoH cago All # Zuzt Pbzt / KOH cago	# 级 + NHs(aq)	→ ① Add drop by drop until in excess
limited	$A ^{3+} + 30H^{-} \rightarrow A (0H)_3$ $Zn^{2+} + 20H^{-} \rightarrow Zn(0H)_2$ $Pb^{2+} + 20H^{-} \rightarrow Pb(0H)_2$ © forms white ppt	ZN ²⁺ +2OH ⁻ → Zn(OH) ₂ } ② forms white ppt Cu ²⁺ +2OH ⁻ → Cu(OH) ₂ } ③ forms <u>blue</u> ppt	
excess	A1(OH) ₃ +OH \rightarrow A1(OH) ₄ Zn(OH) ₂ +2OH \rightarrow Zn(OH) ₄ ² Pb(OH) ₂ +2OH \rightarrow Pb(OH) ₄ ² 3 again	$Zn(0H)_2 + 4NH_5 \rightarrow Zn(NH_5)_4^{2+} + 20H^-$ 23 then dis $Cu(0H)_2 + 4NH_5 \rightarrow Cu(NH_5)_4^{2+} + 20H^-$ 3 organn Coulomb $Coulomb $ $Coulo$	solve

2 Distinguishing tests

- 大部分问题也不会用到 redissolve 的方法
 - > eg. Zn(NO3)2, Pb(NO3)2
 - -> Add HClago, Pb2+ + 2Cl -> PbClz
- 用 redissolve 的例题
 - > eg. Zn(0H)2, A1(0H)3
 - -> Add both solids into NH3 (ag) respectively.
 - → AI(OH)3 does not dissolve, Zn(OH)2 dissolves.
 - \rightarrow ZN(OH)₂ + 4NH₃ \rightarrow ZN(NH₃)₄^{2†} + 2OH⁻