

32

MAXWELL'S EQUATIONS; MAGNETISM OF MATTER

32-1 WHAT IS PHYSICS?

This chapter reveals some of the breadth of physics because it ranges from the basic science of electric and magnetic fields to the applied science and engineering of magnetic materials. First, we conclude our basic discussion of electric and magnetic fields, finding that most of the physics principles in the last 11 chapters can be summarized in only *four* equations, known as Maxwell's equations.

Second, we examine the science and engineering of magnetic materials. The careers of many scientists and engineers are focused on understanding why some materials are magnetic and others are not and on how existing magnetic materials can be improved. These researchers wonder why Earth has a magnetic field but you do not. They find countless applications for inexpensive magnetic materials in cars, kitchens, offices, and hospitals, and magnetic materials often show up in unexpected ways. For example, if you have a tattoo (Fig. 32-1) and undergo an MRI (magnetic resonance imaging) scan, the large magnetic field used in the scan may noticeably tug on your tattooed skin because some tattoo inks contain magnetic particles. In another example, some breakfast cereals are advertised as being "iron fortified" because they contain small bits of iron for you to ingest. Because these iron bits are magnetic, you can collect them by passing a magnet over a slurry of water and cereal.

Our first step here is to revisit Gauss' law, but this time for magnetic fields.



Fig. 32-1 Some of the inks used for tattoos contain magnetic particles.
(Oliver Stewie/Getty Images, Inc.)

32-2 Gauss' Law for Magnetic Fields

Figure 32-2 shows iron powder that has been sprinkled onto a transparent sheet placed above a bar magnet. The powder grains, trying to align themselves with the magnet's magnetic field, have fallen into a pattern that reveals the field. One end of the magnet is a *source* of the field (the field lines diverge from it) and the other end is a *sink* of the field (the field lines converge toward it). By convention, we call the source the *north pole* of the magnet and the sink the *south pole*, and we say that the magnet, with its two poles, is an example of a **magnetic dipole**.



Fig. 32-2 A bar magnet is a magnetic dipole. The iron filings suggest the magnetic field lines. (Colored light fills the background.) (Runk/Schoenberger/Grant Heilman Photography)

Suppose we break a bar magnet into pieces the way we can break a piece of chalk (Fig. 32-3). We should, it seems, be able to isolate a single magnetic pole, called a *magnetic monopole*. However, we cannot—not even if we break the magnet down to its individual atoms and then to its electrons and nuclei. Each fragment has a north pole and a south pole. Thus:



The simplest magnetic structure that can exist is a magnetic dipole. Magnetic monopoles do not exist (as far as we know).

Gauss' law for magnetic fields is a formal way of saying that magnetic monopoles do not exist. The law asserts that the net magnetic flux Φ_B through any closed Gaussian surface is zero:

$$\Phi_B = \oint \vec{B} \cdot d\vec{A} = 0 \quad (\text{Gauss' law for magnetic fields}). \quad (32-1)$$

Contrast this with Gauss' law for electric fields,

$$\Phi_E = \oint \vec{E} \cdot d\vec{A} = \frac{q_{\text{enc}}}{\epsilon_0} \quad (\text{Gauss' law for electric fields}).$$

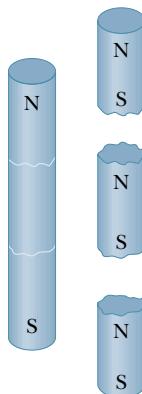


Fig. 32-3 If you break a magnet, each fragment becomes a separate magnet, with its own north and south poles.

In both equations, the integral is taken over a *closed* Gaussian surface. Gauss' law for electric fields says that this integral (the net electric flux through the surface) is proportional to the net electric charge q_{enc} enclosed by the surface. Gauss' law for magnetic fields says that there can be no net magnetic flux through the surface because there can be no net "magnetic charge" (individual magnetic poles) enclosed by the surface. The simplest magnetic structure that can exist and thus be enclosed by a Gaussian surface is a dipole, which consists of both a source and a sink for the field lines. Thus, there must always be as much magnetic flux into the surface as out of it, and the net magnetic flux must always be zero.

Gauss' law for magnetic fields holds for structures more complicated than a magnetic dipole, and it holds even if the Gaussian surface does not enclose the entire structure. Gaussian surface II near the bar magnet of Fig. 32-4 encloses no poles, and we can easily conclude that the net magnetic flux through it is zero. Gaussian surface I is more difficult. It may seem to enclose only the north pole of the magnet because it encloses the label N and not the label S. However, a south pole must be associated with the lower boundary of the surface because magnetic field lines enter the surface there. (The enclosed section is like one piece of the broken bar magnet in Fig. 32-3.) Thus, Gaussian surface I encloses a magnetic dipole, and the net flux through the surface is zero.



CHECKPOINT 1

The figure here shows four closed surfaces with flat top and bottom faces and curved sides. The table gives the areas A of the faces and the magnitudes B of the uniform and perpendicular magnetic fields through those faces; the units of A and B are arbitrary but consistent. Rank the surfaces according to the magnitudes of the magnetic flux through their curved sides, greatest first.

Surface	A_{top}	B_{top}	A_{bot}	B_{bot}
a	2	6, outward	4	3, inward
b	2	1, inward	4	2, inward
c	2	6, inward	2	8, outward
d	2	3, outward	3	2, outward

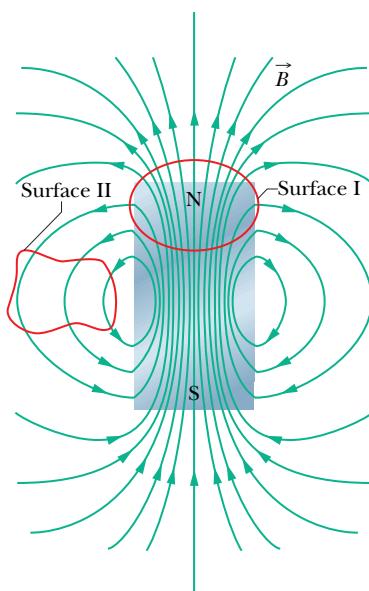
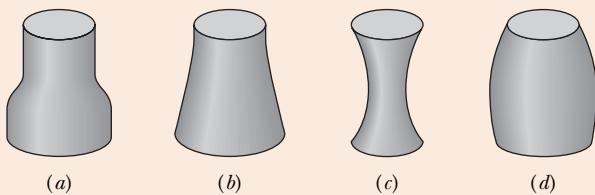


Fig. 32-4 The field lines for the magnetic field \vec{B} of a short bar magnet. The red curves represent cross sections of closed, three-dimensional Gaussian surfaces.

32-3 Induced Magnetic Fields

In Chapter 30 you saw that a changing magnetic flux induces an electric field, and we ended up with Faraday's law of induction in the form

$$\oint \vec{E} \cdot d\vec{s} = -\frac{d\Phi_B}{dt} \quad (\text{Faraday's law of induction}). \quad (32-2)$$

Here \vec{E} is the electric field induced along a closed loop by the changing magnetic flux Φ_B encircled by that loop. Because symmetry is often so powerful in physics, we should be tempted to ask whether induction can occur in the opposite sense; that is, can a changing electric flux induce a magnetic field?

The answer is that it can; furthermore, the equation governing the induction of a magnetic field is almost symmetric with Eq. 32-2. We often call it Maxwell's law of induction after James Clerk Maxwell, and we write it as

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt} \quad (\text{Maxwell's law of induction}). \quad (32-3)$$

Here \vec{B} is the magnetic field induced along a closed loop by the changing electric flux Φ_E in the region encircled by that loop.

As an example of this sort of induction, we consider the charging of a parallel-plate capacitor with circular plates. (Although we shall focus on this arrangement,

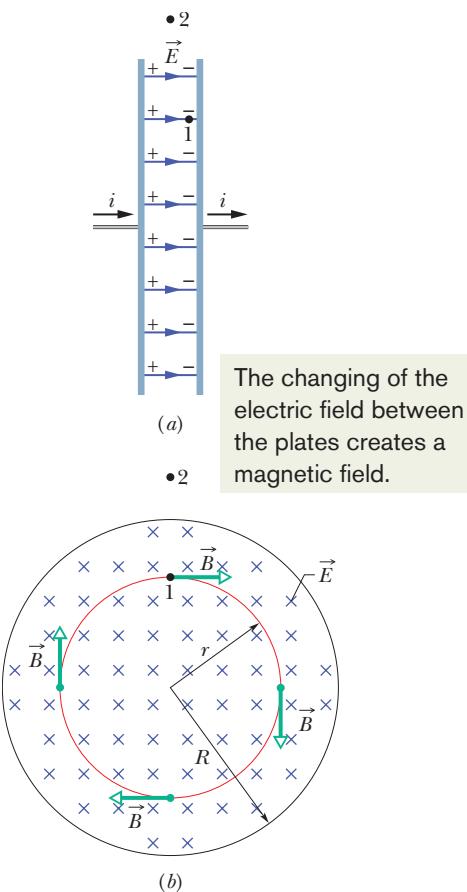


Fig. 32-5 (a) A circular parallel-plate capacitor, shown in side view, is being charged by a constant current i . (b) A view from within the capacitor, looking toward the plate at the right in (a). The electric field \vec{E} is uniform, is directed into the page (toward the plate), and grows in magnitude as the charge on the capacitor increases. The magnetic field \vec{B} induced by this changing electric field is shown at four points on a circle with a radius r less than the plate radius R .

a changing electric flux will always induce a magnetic field whenever it occurs.) We assume that the charge on our capacitor (Fig. 32-5a) is being increased at a steady rate by a constant current i in the connecting wires. Then the electric field magnitude between the plates must also be increasing at a steady rate.

Figure 32-5b is a view of the right-hand plate of Fig. 32-5a from between the plates. The electric field is directed into the page. Let us consider a circular loop through point 1 in Figs. 32-5a and b, a loop that is concentric with the capacitor plates and has a radius smaller than that of the plates. Because the electric field through the loop is changing, the electric flux through the loop must also be changing. According to Eq. 32-3, this changing electric flux induces a magnetic field around the loop.

Experiment proves that a magnetic field \vec{B} is indeed induced around such a loop, directed as shown. This magnetic field has the same magnitude at every point around the loop and thus has circular symmetry about the *central axis* of the capacitor plates (the axis extending from one plate center to the other).

If we now consider a larger loop—say, through point 2 outside the plates in Figs. 32-5a and b—we find that a magnetic field is induced around that loop as well. Thus, while the electric field is changing, magnetic fields are induced between the plates, both inside and outside the gap. When the electric field stops changing, these induced magnetic fields disappear.

Although Eq. 32-3 is similar to Eq. 32-2, the equations differ in two ways. First, Eq. 32-3 has the two extra symbols μ_0 and ϵ_0 , but they appear only because we employ SI units. Second, Eq. 32-3 lacks the minus sign of Eq. 32-2, meaning that the induced electric field \vec{E} and the induced magnetic field \vec{B} have opposite directions when they are produced in otherwise similar situations. To see this opposition, examine Fig. 32-6, in which an increasing magnetic field \vec{B} , directed into the page, induces an electric field \vec{E} . The induced field \vec{E} is counter-clockwise, opposite the induced magnetic field \vec{B} in Fig. 32-5b.

Ampere–Maxwell Law

Now recall that the left side of Eq. 32-3, the integral of the dot product $\vec{B} \cdot d\vec{s}$ around a closed loop, appears in another equation—namely, Ampere's law:

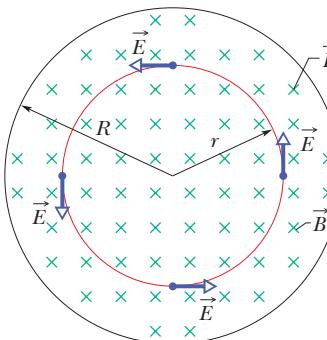
$$\oint \vec{B} \cdot d\vec{s} = \mu_0 i_{\text{enc}} \quad (\text{Ampere's law}), \quad (32-4)$$

where i_{enc} is the current encircled by the closed loop. Thus, our two equations that specify the magnetic field \vec{B} produced by means other than a magnetic material (that is, by a current and by a changing electric field) give the field in exactly the same form. We can combine the two equations into the single equation

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt} + \mu_0 i_{\text{enc}} \quad (\text{Ampere–Maxwell law}). \quad (32-5)$$

The induced \vec{E} direction here is opposite the induced \vec{B} direction in the preceding figure.

Fig. 32-6 A uniform magnetic field \vec{B} in a circular region. The field, directed into the page, is increasing in magnitude. The electric field \vec{E} induced by the changing magnetic field is shown at four points on a circle concentric with the circular region. Compare this situation with that of Fig. 32-5b.



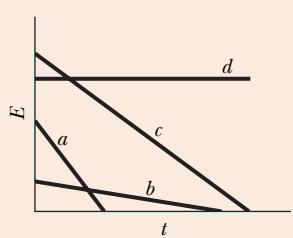
32-3 INDUCED MAGNETIC FIELDS

865

When there is a current but no change in electric flux (such as with a wire carrying a constant current), the first term on the right side of Eq. 32-5 is zero, and so Eq. 32-5 reduces to Eq. 32-4, Ampere's law. When there is a change in electric flux but no current (such as inside or outside the gap of a charging capacitor), the second term on the right side of Eq. 32-5 is zero, and so Eq. 32-5 reduces to Eq. 32-3, Maxwell's law of induction.

**CHECKPOINT 2**

The figure shows graphs of the electric field magnitude E versus time t for four uniform electric fields, all contained within identical circular regions as in Fig. 32-5b. Rank the fields according to the magnitudes of the magnetic fields they induce at the edge of the region, greatest first.

**Sample Problem****Magnetic field induced by changing electric field**

A parallel-plate capacitor with circular plates of radius R is being charged as in Fig. 32-5a.

- (a) Derive an expression for the magnetic field at radius r for the case $r \leq R$.

KEY IDEAS

A magnetic field can be set up by a current and by induction due to a changing electric flux; both effects are included in Eq. 32-5. There is no current between the capacitor plates of Fig. 32-5, but the electric flux there is changing. Thus, Eq. 32-5 reduces to

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt}. \quad (32-6)$$

We shall separately evaluate the left and right sides of this equation.

Left side of Eq. 32-6: We choose a circular Amperian loop with a radius $r \leq R$ as shown in Fig. 32-5b because we want to evaluate the magnetic field for $r \leq R$ —that is, inside the capacitor. The magnetic field \vec{B} at all points along the loop is tangent to the loop, as is the path element $d\vec{s}$. Thus, \vec{B} and $d\vec{s}$ are either parallel or antiparallel at each point of the loop. For simplicity, assume they are parallel (the choice does not alter our outcome here). Then

$$\oint \vec{B} \cdot d\vec{s} = \oint B ds \cos 0^\circ = \oint B ds.$$

Due to the circular symmetry of the plates, we can also assume that \vec{B} has the same magnitude at every point around the loop. Thus, B can be taken outside the integral on the right side of the above equation. The integral that remains is $\oint ds$, which simply gives the circumference $2\pi r$ of the loop. The left side of Eq. 32-6 is then $(B)(2\pi r)$.

Right side of Eq. 32-6: We assume that the electric field \vec{E} is uniform between the capacitor plates and directed perpendicular to the plates. Then the electric flux Φ_E through the Amperian loop is EA , where A is the area encircled by the loop within the electric field. Thus, the right side of Eq. 32-6 is $\mu_0 \epsilon_0 d(EA)/dt$.

Combining results: Substituting our results for the left and right sides into Eq. 32-6, we get

$$(B)(2\pi r) = \mu_0 \epsilon_0 A \frac{dE}{dt}.$$

Because A is a constant, we write $d(EA)$ as $A dE$; so we have

$$(B)(2\pi r) = \mu_0 \epsilon_0 A \frac{dE}{dt}. \quad (32-7)$$

The area A that is encircled by the Amperian loop within the electric field is the *full* area πr^2 of the loop because the loop's radius r is less than (or equal to) the plate radius R . Substituting πr^2 for A in Eq. 32-7 leads to, for $r \leq R$,

$$B = \frac{\mu_0 \epsilon_0 r}{2} \frac{dE}{dt}. \quad (\text{Answer}) \quad (32-8)$$

This equation tells us that, inside the capacitor, B increases linearly with increased radial distance r , from 0 at the central axis to a maximum value at plate radius R .

- (b) Evaluate the field magnitude B for $r = R/5 = 11.0$ mm and $dE/dt = 1.50 \times 10^{12}$ V/m · s.

Calculation: From the answer to (a), we have

$$\begin{aligned} B &= \frac{1}{2} \mu_0 \epsilon_0 r \frac{dE}{dt} \\ &= \frac{1}{2} (4\pi \times 10^{-7} \text{ T} \cdot \text{m}/\text{A})(8.85 \times 10^{-12} \text{ C}^2/\text{N} \cdot \text{m}^2) \\ &\quad \times (11.0 \times 10^{-3} \text{ m})(1.50 \times 10^{12} \text{ V/m} \cdot \text{s}) \\ &= 9.18 \times 10^{-8} \text{ T}. \end{aligned} \quad (\text{Answer})$$

- (c) Derive an expression for the induced magnetic field for the case $r \geq R$.

Calculation: Our procedure is the same as in (a) except we now use an Amperian loop with a radius r that is greater than the plate radius R , to evaluate B outside the capacitor. Evaluating the left and right sides of Eq. 32-6 again leads to Eq. 32-7. However, we then need this subtle point: The electric field exists only between the plates, not outside the plates. Thus, the area A that is encircled by the Amperian loop in the electric field is *not* the full area πr^2 of the loop. Rather, A is only the plate area πR^2 .

Substituting πR^2 for A in Eq. 32-7 and solving the result for B give us, for $r \geq R$,

$$B = \frac{\mu_0 \epsilon_0 R^2}{2r} \frac{dE}{dt}. \quad (\text{Answer}) \quad (32-9)$$



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This equation tells us that, outside the capacitor, B decreases with increased radial distance r , from a maximum value at the plate edges (where $r = R$). By substituting $r = R$ into Eqs. 32-8 and 32-9, you can show that these equations are consistent; that is, they give the same maximum value of B at the plate radius.

The magnitude of the induced magnetic field calculated in (b) is so small that it can scarcely be measured with simple apparatus. This is in sharp contrast to the magnitudes of induced electric fields (Faraday's law), which can be measured easily. This experimental difference exists partly because induced emfs can easily be multiplied by using a coil of many turns. No technique of comparable simplicity exists for multiplying induced magnetic fields. In any case, the experiment suggested by this sample problem has been done, and the presence of the induced magnetic fields has been verified quantitatively.

32-4 Displacement Current

If you compare the two terms on the right side of Eq. 32-5, you will see that the product $\epsilon_0(d\Phi_E/dt)$ must have the dimension of a current. In fact, that product has been treated as being a fictitious current called the **displacement current** i_d :

$$i_d = \epsilon_0 \frac{d\Phi_E}{dt} \quad (\text{displacement current}). \quad (32-10)$$

"Displacement" is poorly chosen in that nothing is being displaced, but we are stuck with the word. Nevertheless, we can now rewrite Eq. 32-5 as

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 i_{d,\text{enc}} + \mu_0 i_{\text{enc}} \quad (\text{Ampere-Maxwell law}), \quad (32-11)$$

in which $i_{d,\text{enc}}$ is the displacement current that is encircled by the integration loop.

Let us again focus on a charging capacitor with circular plates, as in Fig. 32-7a. The real current i that is charging the plates changes the electric field \vec{E} between the plates. The fictitious displacement current i_d between the plates is associated with that changing field \vec{E} . Let us relate these two currents.

The charge q on the plates at any time is related to the magnitude E of the field between the plates at that time by Eq. 25-4:

$$q = \epsilon_0 A E, \quad (32-12)$$

in which A is the plate area. To get the real current i , we differentiate Eq. 32-12 with respect to time, finding

$$\frac{dq}{dt} = i = \epsilon_0 A \frac{dE}{dt}. \quad (32-13)$$

To get the displacement current i_d , we can use Eq. 32-10. Assuming that the electric field \vec{E} between the two plates is uniform (we neglect any fringing), we can replace the electric flux Φ_E in that equation with EA . Then Eq. 32-10

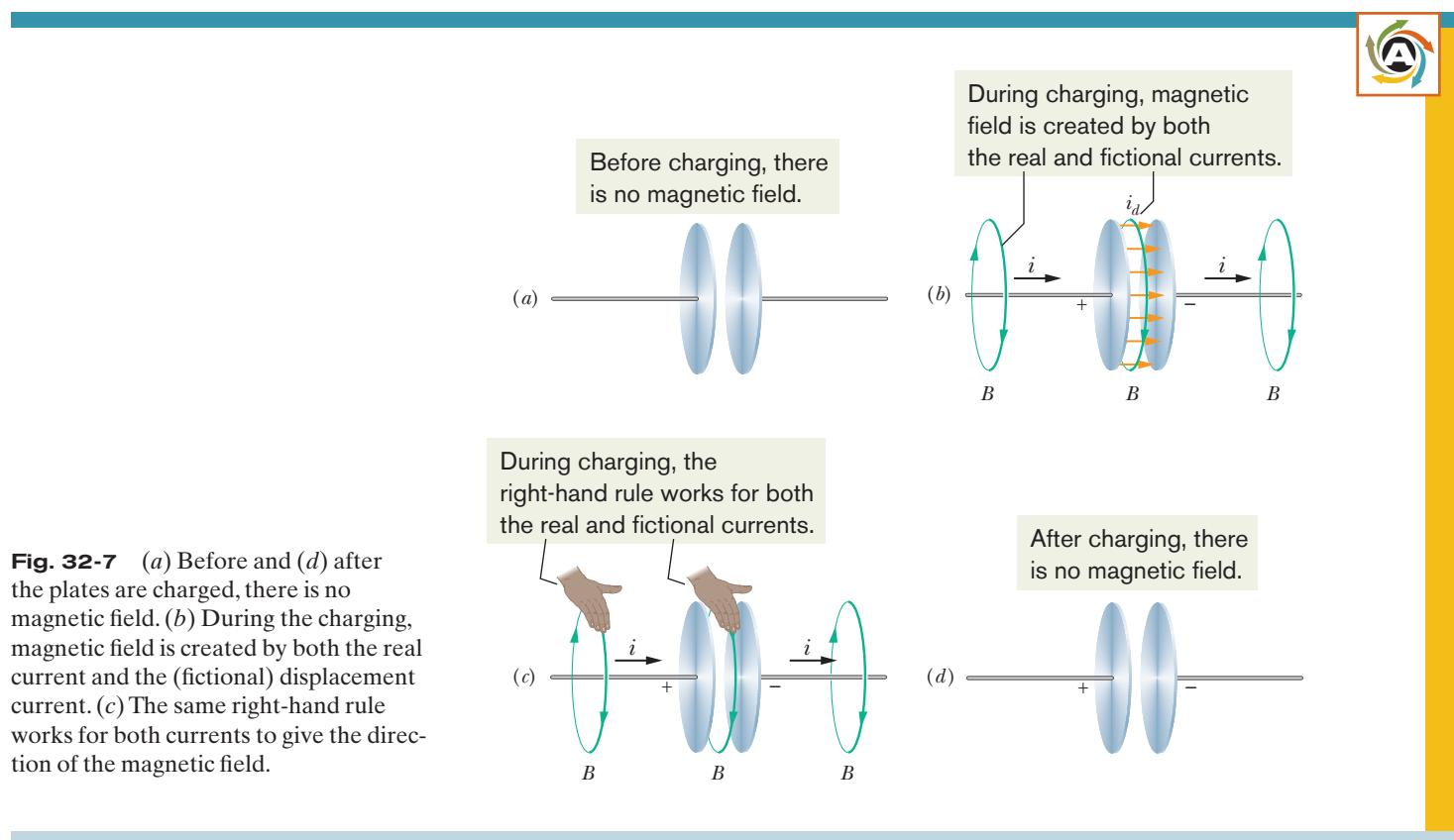


Fig. 32-7 (a) Before and (d) after the plates are charged, there is no magnetic field. (b) During the charging, magnetic field is created by both the real current and the (fictional) displacement current. (c) The same right-hand rule works for both currents to give the direction of the magnetic field.

becomes

$$i_d = \epsilon_0 \frac{d\Phi_E}{dt} = \epsilon_0 \frac{d(EA)}{dt} = \epsilon_0 A \frac{dE}{dt}. \quad (32-14)$$

Comparing Eqs. 32-13 and 32-14, we see that the real current i charging the capacitor and the fictitious displacement current i_d between the plates have the same magnitude:

$$i_d = i \quad (\text{displacement current in a capacitor}). \quad (32-15)$$

Thus, we can consider the fictitious displacement current i_d to be simply a continuation of the real current i from one plate, across the capacitor gap, to the other plate. Because the electric field is uniformly spread over the plates, the same is true of this fictitious displacement current i_d , as suggested by the spread of current arrows in Fig. 32-7b. Although no charge actually moves across the gap between the plates, the idea of the fictitious current i_d can help us to quickly find the direction and magnitude of an induced magnetic field, as follows.

Finding the Induced Magnetic Field

In Chapter 29 we found the direction of the magnetic field produced by a real current i by using the right-hand rule of Fig. 29-4. We can apply the same rule to find the direction of an induced magnetic field produced by a fictitious displacement current i_d , as is shown in the center of Fig. 32-7c for a capacitor.

We can also use i_d to find the magnitude of the magnetic field induced by a charging capacitor with parallel circular plates of radius R . We simply consider the space between the plates to be an imaginary circular wire of radius R carrying the imaginary current i_d . Then, from Eq. 29-20, the magnitude of the magnetic

field at a point inside the capacitor at radius r from the center is

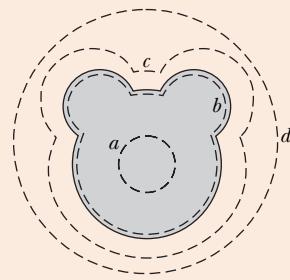
$$B = \left(\frac{\mu_0 i_d}{2\pi R^2} \right) r \quad (\text{inside a circular capacitor}). \quad (32-16)$$

Similarly, from Eq. 29-17, the magnitude of the magnetic field at a point outside the capacitor at radius r is

$$B = \frac{\mu_0 i_d}{2\pi r} \quad (\text{outside a circular capacitor}). \quad (32-17)$$

CHECKPOINT 3

The figure is a view of one plate of a parallel-plate capacitor from within the capacitor. The dashed lines show four integration paths (path b follows the edge of the plate). Rank the paths according to the magnitude of $\oint \vec{B} \cdot d\vec{s}$ along the paths during the discharging of the capacitor, greatest first.



Sample Problem

Treating a changing electric field as a displacement current

A circular parallel-plate capacitor with plate radius R is being charged with a current i . This gives us

- (a) Between the plates, what is the magnitude of $\oint \vec{B} \cdot d\vec{s}$, in terms of μ_0 and i , at a radius $r = R/5$ from their center?

KEY IDEA

A magnetic field can be set up by a current and by induction due to a changing electric flux (Eq. 32-5). Between the plates in Fig. 32-5, the current is zero and we can account for the changing electric flux with a fictitious displacement current i_d . Then integral $\oint \vec{B} \cdot d\vec{s}$ is given by Eq. 32-11, but because there is no real current i between the capacitor plates, the equation reduces to

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 i_{d,\text{enc}}. \quad (32-18)$$

Calculations: Because we want to evaluate $\oint \vec{B} \cdot d\vec{s}$ at radius $r = R/5$ (within the capacitor), the integration loop encircles only a portion $i_{d,\text{enc}}$ of the total displacement current i_d . Let's assume that i_d is uniformly spread over the full plate area. Then the portion of the displacement current encircled by the loop is proportional to the area encircled by the loop:

$$\frac{\left(\begin{array}{c} \text{encircled displacement} \\ \text{current } i_{d,\text{enc}} \end{array} \right)}{\left(\begin{array}{c} \text{total displacement} \\ \text{current } i_d \end{array} \right)} = \frac{\text{encircled area } \pi r^2}{\text{full plate area } \pi R^2}.$$

$$i_{d,\text{enc}} = i_d \frac{\pi r^2}{\pi R^2}.$$

Substituting this into Eq. 32-18, we obtain

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 i_d \frac{\pi r^2}{\pi R^2}. \quad (32-19)$$

Now substituting $i_d = i$ (from Eq. 32-15) and $r = R/5$ into Eq. 32-19 leads to

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 i \frac{(R/5)^2}{R^2} = \frac{\mu_0 i}{25}. \quad (\text{Answer})$$

- (b) In terms of the maximum induced magnetic field, what is the magnitude of the magnetic field induced at $r = R/5$, inside the capacitor?

KEY IDEA

Because the capacitor has parallel circular plates, we can treat the space between the plates as an imaginary wire of radius R carrying the imaginary current i_d . Then we can use Eq. 32-16 to find the induced magnetic field magnitude B at any point inside the capacitor.

- Calculations:** At $r = R/5$, Eq. 32-16 yields

$$B = \left(\frac{\mu_0 i_d}{2\pi R^2} \right) r = \frac{\mu_0 i_d (R/5)}{2\pi R^2} = \frac{\mu_0 i_d}{10\pi R}. \quad (32-20)$$

From Eq. 32-16, the maximum field magnitude B_{\max} within the capacitor occurs at $r = R$. It is

$$B_{\max} = \left(\frac{\mu_0 i_d}{2\pi R^2} \right) R = \frac{\mu_0 i_d}{2\pi R}. \quad (32-21)$$

Dividing Eq. 32-20 by Eq. 32-21 and rearranging the result, we find that the field magnitude at $r = R/5$ is

$$B = \frac{1}{5} B_{\max}. \quad (\text{Answer})$$

We should be able to obtain this result with a little reasoning and less work. Equation 32-16 tells us that inside the capacitor, B increases linearly with r . Therefore, a point $\frac{1}{5}$ the distance out to the full radius R of the plates, where B_{\max} occurs, should have a field B that is $\frac{1}{5} B_{\max}$.



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32-5 Maxwell's Equations

Equation 32-5 is the last of the four fundamental equations of electromagnetism, called *Maxwell's equations* and displayed in Table 32-1. These four equations explain a diverse range of phenomena, from why a compass needle points north to why a car starts when you turn the ignition key. They are the basis for the functioning of such electromagnetic devices as electric motors, television transmitters and receivers, telephones, fax machines, radar, and microwave ovens.

Maxwell's equations are the basis from which many of the equations you have seen since Chapter 21 can be derived. They are also the basis of many of the equations you will see in Chapters 33 through 36 concerning optics.

Table 32-1

Maxwell's Equations^a

Name	Equation	
Gauss' law for electricity	$\oint \vec{E} \cdot d\vec{A} = q_{\text{enc}}/\epsilon_0$	Relates net electric flux to net enclosed electric charge
Gauss' law for magnetism	$\oint \vec{B} \cdot d\vec{A} = 0$	Relates net magnetic flux to net enclosed magnetic charge
Faraday's law	$\oint \vec{E} \cdot d\vec{s} = -\frac{d\Phi_B}{dt}$	Relates induced electric field to changing magnetic flux
Ampere–Maxwell law	$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt} + \mu_0 i_{\text{enc}}$	Relates induced magnetic field to changing electric flux and to current

^aWritten on the assumption that no dielectric or magnetic materials are present.

32-6 Magnets

The first known magnets were *lodestones*, which are stones that have been *magnetized* (made magnetic) naturally. When the ancient Greeks and ancient Chinese discovered these rare stones, they were amused by the stones' ability to attract metal over a short distance, as if by magic. Only much later did they learn to use lodestones (and artificially magnetized pieces of iron) in compasses to determine direction.

Today, magnets and magnetic materials are ubiquitous. Their magnetic properties can be traced to their atoms and electrons. In fact, the inexpensive magnet you might use to hold a note on the refrigerator door is a direct result of the quantum physics taking place in the atomic and subatomic material within the magnet. Before we explore some of this physics, let's briefly discuss the largest magnet we commonly use—namely, Earth itself.

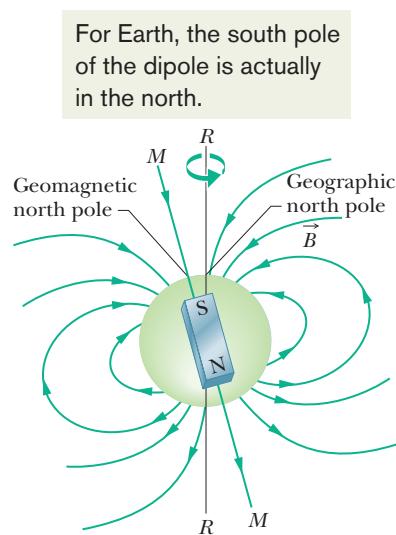


Fig. 32-8 Earth's magnetic field represented as a dipole field. The dipole axis MM makes an angle of 11.5° with Earth's rotational axis RR . The south pole of the dipole is in Earth's Northern Hemisphere.

The Magnetism of Earth

Earth is a huge magnet; for points near Earth's surface, its magnetic field can be approximated as the field of a huge bar magnet—a magnetic dipole—that straddles the center of the planet. Figure 32-8 is an idealized symmetric depiction of the dipole field, without the distortion caused by passing charged particles from the Sun.

Because Earth's magnetic field is that of a magnetic dipole, a magnetic dipole moment $\vec{\mu}$ is associated with the field. For the idealized field of Fig. 32-8, the magnitude of $\vec{\mu}$ is $8.0 \times 10^{22} \text{ J/T}$ and the direction of $\vec{\mu}$ makes an angle of 11.5° with the rotation axis (RR) of Earth. The *dipole axis* (MM in Fig. 32-8) lies along $\vec{\mu}$ and intersects Earth's surface at the *geomagnetic north pole* off the northwest coast of Greenland and the *geomagnetic south pole* in Antarctica. The lines of the magnetic field \vec{B} generally emerge in the Southern Hemisphere and reenter Earth in the Northern Hemisphere. Thus, the magnetic pole that is in Earth's Northern Hemisphere and known as a “north magnetic pole” is *really the south pole of Earth's magnetic dipole*.

The direction of the magnetic field at any location on Earth's surface is commonly specified in terms of two angles. The **field declination** is the angle (left or right) between geographic north (which is toward 90° latitude) and the horizontal component of the field. The **field inclination** is the angle (up or down) between a horizontal plane and the field's direction.

Magnetometers measure these angles and determine the field with much precision. However, you can do reasonably well with just a *compass* and a *dip meter*. A compass is simply a needle-shaped magnet that is mounted so it can rotate freely about a vertical axis. When it is held in a horizontal plane, the north-pole end of the needle points, generally, toward the geomagnetic north pole (really a south magnetic pole, remember). The angle between the needle and geographic north is the field declination. A dip meter is a similar magnet that can rotate freely about a horizontal axis. When its vertical plane of rotation is aligned with the direction of the compass, the angle between the meter's needle and the horizontal is the field inclination.

At any point on Earth's surface, the measured magnetic field may differ appreciably, in both magnitude and direction, from the idealized dipole field of Fig. 32-8. In fact, the point where the field is actually perpendicular to Earth's surface and inward is not located at the geomagnetic north pole off Greenland as we would expect; instead, this so-called *dip north pole* is located in the Queen Elizabeth Islands in northern Canada, far from Greenland.

In addition, the field observed at any location on the surface of Earth varies with time, by measurable amounts over a period of a few years and by substantial amounts over, say, 100 years. For example, between 1580 and 1820 the direction indicated by compass needles in London changed by 35° .

In spite of these local variations, the average dipole field changes only slowly over such relatively short time periods. Variations over longer periods can be studied by measuring the weak magnetism of the ocean floor on either side of the Mid-Atlantic Ridge (Fig. 32-9). This floor has been formed by molten magma

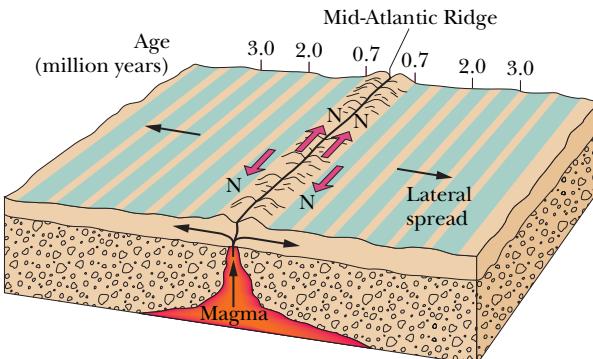


Fig. 32-9 A magnetic profile of the seafloor on either side of the Mid-Atlantic Ridge. The seafloor, extruded through the ridge and spreading out as part of the tectonic drift system, displays a record of the past magnetic history of Earth's core. The direction of the magnetic field produced by the core reverses about every million years.

that oozed up through the ridge from Earth's interior, solidified, and was pulled away from the ridge (by the drift of tectonic plates) at the rate of a few centimeters per year. As the magma solidified, it became weakly magnetized with its magnetic field in the direction of Earth's magnetic field at the time of solidification. Study of this solidified magma across the ocean floor reveals that Earth's field has reversed its *polarity* (directions of the north pole and south pole) about every million years. The reason for the reversals is not known. In fact, the mechanism that produces Earth's magnetic field is only vaguely understood.

32-7 Magnetism and Electrons

Magnetic materials, from lodestones to videotapes, are magnetic because of the electrons within them. We have already seen one way in which electrons can generate a magnetic field: Send them through a wire as an electric current, and their motion produces a magnetic field around the wire. There are two more ways, each involving a magnetic dipole moment that produces a magnetic field in the surrounding space. However, their explanation requires quantum physics that is beyond the physics presented in this book, and so here we shall only outline the results.

Spin Magnetic Dipole Moment

An electron has an intrinsic angular momentum called its **spin angular momentum** (or just **spin**) \vec{S} ; associated with this spin is an intrinsic **spin magnetic dipole moment** $\vec{\mu}_s$. (By *intrinsic*, we mean that \vec{S} and $\vec{\mu}_s$ are basic characteristics of an electron, like its mass and electric charge.) Vectors \vec{S} and $\vec{\mu}_s$ are related by

$$\vec{\mu}_s = -\frac{e}{m} \vec{S} \quad (32-22)$$

in which e is the elementary charge (1.60×10^{-19} C) and m is the mass of an electron (9.11×10^{-31} kg). The minus sign means that $\vec{\mu}_s$ and \vec{S} are oppositely directed.

Spin \vec{S} is different from the angular momenta of Chapter 11 in two respects:

1. Spin \vec{S} itself cannot be measured. However, its component along any axis can be measured.
2. A measured component of \vec{S} is *quantized*, which is a general term that means it is restricted to certain values. A measured component of \vec{S} can have only two values, which differ only in sign.

Let us assume that the component of spin \vec{S} is measured along the z axis of a coordinate system. Then the measured component S_z can have only the two values given by

$$S_z = m_s \frac{h}{2\pi}, \quad \text{for } m_s = \pm \frac{1}{2}, \quad (32-23)$$

where m_s is called the *spin magnetic quantum number* and h ($= 6.63 \times 10^{-34}$ J·s) is the Planck constant, the ubiquitous constant of quantum physics. The signs given in Eq. 32-23 have to do with the direction of S_z along the z axis. When S_z is parallel to the z axis, m_s is $+\frac{1}{2}$ and the electron is said to be *spin up*. When S_z is antiparallel to the z axis, m_s is $-\frac{1}{2}$ and the electron is said to be *spin down*.

The spin magnetic dipole moment $\vec{\mu}_s$ of an electron also cannot be measured; only its component along any axis can be measured, and that component too is quantized, with two possible values of the same magnitude but different signs. We can relate the component $\mu_{s,z}$ measured on the z axis to S_z by rewriting Eq. 32-22 in component form for the z axis as

$$\mu_{s,z} = -\frac{e}{m} S_z.$$

For an electron, the spin is opposite the magnetic dipole moment.

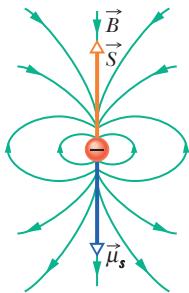


Fig. 32-10 The spin \vec{S} , spin magnetic dipole moment $\vec{\mu}_s$, and magnetic dipole field \vec{B} of an electron represented as a microscopic sphere.

Substituting for S_z from Eq. 32-23 then gives us

$$\mu_{s,z} = \pm \frac{e\hbar}{4\pi m}, \quad (32-24)$$

where the plus and minus signs correspond to $\mu_{s,z}$ being parallel and antiparallel to the z axis, respectively.

The quantity on the right side of Eq. 32-24 is called the *Bohr magneton* μ_B :

$$\mu_B = \frac{e\hbar}{4\pi m} = 9.27 \times 10^{-24} \text{ J/T} \quad (\text{Bohr magneton}). \quad (32-25)$$

Spin magnetic dipole moments of electrons and other elementary particles can be expressed in terms of μ_B . For an electron, the magnitude of the measured z component of $\vec{\mu}_s$ is

$$|\mu_{s,z}| = 1\mu_B. \quad (32-26)$$

(The quantum physics of the electron, called *quantum electrodynamics*, or QED, reveals that $\mu_{s,z}$ is actually slightly greater than $1\mu_B$, but we shall neglect that fact.)

When an electron is placed in an external magnetic field \vec{B}_{ext} , an energy U can be associated with the orientation of the electron's spin magnetic dipole moment $\vec{\mu}_s$ just as an energy can be associated with the orientation of the magnetic dipole moment $\vec{\mu}$ of a current loop placed in \vec{B}_{ext} . From Eq. 28-38, the orientation energy for the electron is

$$U = -\vec{\mu}_s \cdot \vec{B}_{\text{ext}} = -\mu_{s,z} B_{\text{ext}}, \quad (32-27)$$

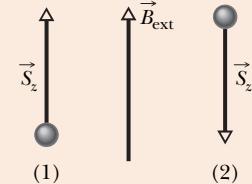
where the z axis is taken to be in the direction of \vec{B}_{ext} .

If we imagine an electron to be a microscopic sphere (which it is not), we can represent the spin \vec{S} , the spin magnetic dipole moment $\vec{\mu}_s$, and the associated magnetic dipole field as in Fig. 32-10. Although we use the word "spin" here, electrons do not spin like tops. How, then, can something have angular momentum without actually rotating? Again, we would need quantum physics to provide the answer.

Protons and neutrons also have an intrinsic angular momentum called spin and an associated intrinsic spin magnetic dipole moment. For a proton those two vectors have the same direction, and for a neutron they have opposite directions. We shall not examine the contributions of these dipole moments to the magnetic fields of atoms because they are about a thousand times smaller than that due to an electron.

CHECKPOINT 4

The figure here shows the spin orientations of two particles in an external magnetic field \vec{B}_{ext} . (a) If the particles are electrons, which spin orientation is at lower energy? (b) If, instead, the particles are protons, which spin orientation is at lower energy?



Orbital Magnetic Dipole Moment

When it is in an atom, an electron has an additional angular momentum called its **orbital angular momentum** \vec{L}_{orb} . Associated with \vec{L}_{orb} is an **orbital magnetic dipole moment** $\vec{\mu}_{\text{orb}}$; the two are related by

$$\vec{\mu}_{\text{orb}} = -\frac{e}{2m} \vec{L}_{\text{orb}}. \quad (32-28)$$

The minus sign means that $\vec{\mu}_{\text{orb}}$ and \vec{L}_{orb} have opposite directions.

Orbital angular momentum \vec{L}_{orb} cannot be measured; only its component along any axis can be measured, and that component is quantized. The component along, say, a z axis can have only the values given by

$$L_{\text{orb},z} = m_\ell \frac{\hbar}{2\pi}, \quad \text{for } m_\ell = 0, \pm 1, \pm 2, \dots, \pm (\text{limit}), \quad (32-29)$$

in which m_ℓ is called the *orbital magnetic quantum number* and “limit” refers to some largest allowed integer value for m_ℓ . The signs in Eq. 32-29 have to do with the direction of $L_{\text{orb},z}$ along the z axis.

The orbital magnetic dipole moment $\vec{\mu}_{\text{orb}}$ of an electron also cannot itself be measured; only its component along an axis can be measured, and that component is quantized. By writing Eq. 32-28 for a component along the same z axis as above and then substituting for $L_{\text{orb},z}$ from Eq. 32-29, we can write the z component $\mu_{\text{orb},z}$ of the orbital magnetic dipole moment as

$$\mu_{\text{orb},z} = -m_\ell \frac{e\hbar}{4\pi m} \quad (32-30)$$

and, in terms of the Bohr magneton, as

$$\mu_{\text{orb},z} = -m_\ell \mu_B. \quad (32-31)$$

When an atom is placed in an external magnetic field \vec{B}_{ext} , an energy U can be associated with the orientation of the orbital magnetic dipole moment of each electron in the atom. Its value is

$$U = -\vec{\mu}_{\text{orb}} \cdot \vec{B}_{\text{ext}} = -\mu_{\text{orb},z} B_{\text{ext}}, \quad (32-32)$$

where the z axis is taken in the direction of \vec{B}_{ext} .

Although we have used the words “orbit” and “orbital” here, electrons do not orbit the nucleus of an atom like planets orbiting the Sun. How can an electron have an orbital angular momentum without orbiting in the common meaning of the term? Once again, this can be explained only with quantum physics.

Loop Model for Electron Orbits

We can obtain Eq. 32-28 with the nonquantum derivation that follows, in which we assume that an electron moves along a circular path with a radius that is much larger than an atomic radius (hence the name “loop model”). However, the derivation does not apply to an electron within an atom (for which we need quantum physics).

We imagine an electron moving at constant speed v in a circular path of radius r , counterclockwise as shown in Fig. 32-11. The motion of the negative charge of the electron is equivalent to a conventional current i (of positive charge) that is clockwise, as also shown in Fig. 32-11. The magnitude of the orbital magnetic dipole moment of such a *current loop* is obtained from Eq. 28-35 with $N = 1$:

$$\mu_{\text{orb}} = iA, \quad (32-33)$$

where A is the area enclosed by the loop. The direction of this magnetic dipole moment is, from the right-hand rule of Fig. 29-21, downward in Fig. 32-11.

To evaluate Eq. 32-33, we need the current i . Current is, generally, the rate at which charge passes some point in a circuit. Here, the charge of magnitude e takes a time $T = 2\pi r/v$ to circle from any point back through that point, so

$$i = \frac{\text{charge}}{\text{time}} = \frac{e}{2\pi r/v}. \quad (32-34)$$

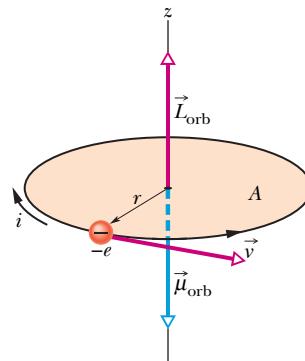


Fig. 32-11 An electron moving at constant speed v in a circular path of radius r that encloses an area A . The electron has an orbital angular momentum \vec{L}_{orb} and an associated orbital magnetic dipole moment $\vec{\mu}_{\text{orb}}$. A clockwise current i (of positive charge) is equivalent to the counterclockwise circulation of the negatively charged electron.

Substituting this and the area $A = \pi r^2$ of the loop into Eq. 32-33 gives us

$$\mu_{\text{orb}} = \frac{e}{2\pi r/v} \pi r^2 = \frac{evr}{2}. \quad (32-35)$$

To find the electron's orbital angular momentum \vec{L}_{orb} , we use Eq. 11-18, $\vec{\ell} = m(\vec{r} \times \vec{v})$. Because \vec{r} and \vec{v} are perpendicular, \vec{L}_{orb} has the magnitude

$$L_{\text{orb}} = mr v \sin 90^\circ = mr v. \quad (32-36)$$

The vector \vec{L}_{orb} is directed upward in Fig. 32-11 (see Fig. 11-12). Combining Eqs. 32-35 and 32-36, generalizing to a vector formulation, and indicating the opposite directions of the vectors with a minus sign yield

$$\vec{\mu}_{\text{orb}} = -\frac{e}{2m} \vec{L}_{\text{orb}},$$

which is Eq. 32-28. Thus, by "classical" (nonquantum) analysis we have obtained the same result, in both magnitude and direction, given by quantum physics. You might wonder, seeing as this derivation gives the correct result for an electron within an atom, why the derivation is invalid for that situation. The answer is that this line of reasoning yields other results that are contradicted by experiments.

Loop Model in a Nonuniform Field

We continue to consider an electron orbit as a current loop, as we did in Fig. 32-11. Now, however, we draw the loop in a nonuniform magnetic field \vec{B}_{ext} as shown in Fig. 32-12a. (This field could be the diverging field near the north pole of the magnet in Fig. 32-4.) We make this change to prepare for the next several sections, in which we shall discuss the forces that act on magnetic materials when the materials are placed in a nonuniform magnetic field. We shall discuss these forces by assuming that the electron orbits in the materials are tiny current loops like that in Fig. 32-12a.

Here we assume that the magnetic field vectors all around the electron's circular path have the same magnitude and form the same angle with the vertical, as shown in Figs. 32-12b and d. We also assume that all the electrons in an atom move either counterclockwise (Fig. 32-12b) or clockwise (Fig. 32-12d). The associated conventional current i around the current loop and the orbital magnetic dipole moment $\vec{\mu}_{\text{orb}}$ produced by i are shown for each direction of motion.

Figures 32-12c and e show diametrically opposite views of a length element $d\vec{L}$ of the loop that has the same direction as i , as seen from the plane of the orbit. Also shown are the field \vec{B}_{ext} and the resulting magnetic force $d\vec{F}$ on $d\vec{L}$. Recall that a current along an element $d\vec{L}$ in a magnetic field \vec{B}_{ext} experiences a magnetic force $d\vec{F}$ as given by Eq. 28-28:

$$d\vec{F} = i d\vec{L} \times \vec{B}_{\text{ext}}. \quad (32-37)$$

On the left side of Fig. 32-12c, Eq. 32-37 tells us that the force $d\vec{F}$ is directed upward and rightward. On the right side, the force $d\vec{F}$ is just as large and is directed upward and leftward. Because their angles are the same, the horizontal components of these two forces cancel and the vertical components add. The same is true at any other two symmetric points on the loop. Thus, the net force on the current loop of Fig. 32-12b must be upward. The same reasoning leads to a downward net force on the loop in Fig. 32-12d. We shall use these two results shortly when we examine the behavior of magnetic materials in nonuniform magnetic fields.

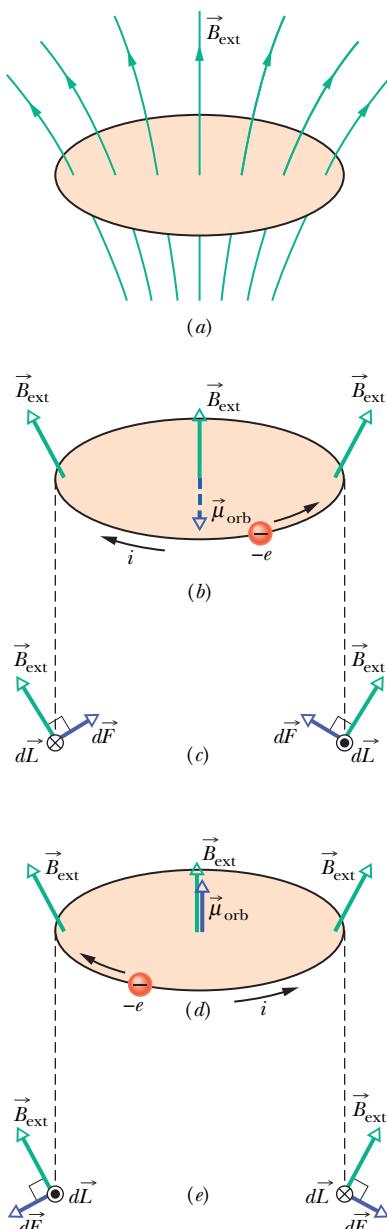


Fig. 32-12 (a) A loop model for an electron orbiting in an atom while in a nonuniform magnetic field \vec{B}_{ext} . (b) Charge $-e$ moves counterclockwise; the associated conventional current i is clockwise. (c) The magnetic forces $d\vec{F}$ on the left and right sides of the loop, as seen from the plane of the loop. The net force on the loop is upward. (d) Charge $-e$ now moves clockwise. (e) The net force on the loop is now downward.

32-8 Magnetic Materials

Each electron in an atom has an orbital magnetic dipole moment and a spin magnetic dipole moment that combine vectorially. The resultant of these two vector quantities combines vectorially with similar resultants for all other electrons in the atom, and the resultant for each atom combines with those for all the other atoms in a sample of a material. If the combination of all these magnetic dipole moments produces a magnetic field, then the material is magnetic. There are three general types of magnetism: diamagnetism, paramagnetism, and ferromagnetism.

- Diamagnetism** is exhibited by all common materials but is so feeble that it is masked if the material also exhibits magnetism of either of the other two types. In diamagnetism, weak magnetic dipole moments are produced in the atoms of the material when the material is placed in an external magnetic field \vec{B}_{ext} ; the combination of all those induced dipole moments gives the material as a whole only a feeble net magnetic field. The dipole moments and thus their net field disappear when \vec{B}_{ext} is removed. The term *diamagnetic material* usually refers to materials that exhibit only diamagnetism.
- Paramagnetism** is exhibited by materials containing transition elements, rare earth elements, and actinide elements (see Appendix G). Each atom of such a material has a permanent resultant magnetic dipole moment, but the moments are randomly oriented in the material and the material as a whole lacks a net magnetic field. However, an external magnetic field \vec{B}_{ext} can partially align the atomic magnetic dipole moments to give the material a net magnetic field. The alignment and thus its field disappear when \vec{B}_{ext} is removed. The term *paramagnetic material* usually refers to materials that exhibit primarily paramagnetism.
- Ferromagnetism** is a property of iron, nickel, and certain other elements (and of compounds and alloys of these elements). Some of the electrons in these materials have their resultant magnetic dipole moments aligned, which produces regions with strong magnetic dipole moments. An external field \vec{B}_{ext} can then align the magnetic moments of such regions, producing a strong magnetic field for a sample of the material; the field partially persists when \vec{B}_{ext} is removed. We usually use the terms *ferromagnetic material* and *magnetic material* to refer to materials that exhibit primarily ferromagnetism.

The next three sections explore these three types of magnetism.

32-9 Diamagnetism

We cannot yet discuss the quantum physical explanation of diamagnetism, but we can provide a classical explanation with the loop model of Figs. 32-11 and 32-12. To begin, we assume that in an atom of a diamagnetic material each electron can orbit only clockwise as in Fig. 32-12d or counterclockwise as in Fig. 32-12b. To account for the lack of magnetism in the absence of an external magnetic field \vec{B}_{ext} , we assume the atom lacks a net magnetic dipole moment. This implies that before \vec{B}_{ext} is applied, the number of electrons orbiting in one direction is the same as that orbiting in the opposite direction, with the result that the net upward magnetic dipole moment of the atom equals the net downward magnetic dipole moment.

Now let's turn on the nonuniform field \vec{B}_{ext} of Fig. 32-12a, in which \vec{B}_{ext} is directed upward but is diverging (the magnetic field lines are diverging). We could do this by increasing the current through an electromagnet or by moving



Fig. 32-13 An overhead view of a frog that is being levitated in a magnetic field produced by current in a vertical solenoid below the frog. (Courtesy A. K. Gein, High Field Magnet Laboratory, University of Nijmegen, The Netherlands)

the north pole of a bar magnet closer to, and below, the orbits. As the magnitude of \vec{B}_{ext} increases from zero to its final maximum, steady-state value, a clockwise electric field is induced around each electron's orbital loop according to Faraday's law and Lenz's law. Let us see how this induced electric field affects the orbiting electrons in Figs. 32-12b and d.

In Fig. 32-12b, the counterclockwise electron is accelerated by the clockwise electric field. Thus, as the magnetic field \vec{B}_{ext} increases to its maximum value, the electron speed increases to a maximum value. This means that the associated conventional current i and the downward magnetic dipole moment $\vec{\mu}$ due to i also *increase*.

In Fig. 32-12d, the clockwise electron is decelerated by the clockwise electric field. Thus, here, the electron speed, the associated current i , and the upward magnetic dipole moment $\vec{\mu}$ due to i all *decrease*. By turning on field \vec{B}_{ext} , we have given the atom a *net* magnetic dipole moment that is downward. This would also be so if the magnetic field were uniform.

The nonuniformity of field \vec{B}_{ext} also affects the atom. Because the current i in Fig. 32-12b increases, the upward magnetic forces $d\vec{F}$ in Fig. 32-12c also increase, as does the net upward force on the current loop. Because current i in Fig. 32-12d decreases, the downward magnetic forces $d\vec{F}$ in Fig. 32-12e also decrease, as does the net downward force on the current loop. Thus, by turning on the *nonuniform* field \vec{B}_{ext} , we have produced a net force on the atom; moreover, that force is directed *away* from the region of greater magnetic field.

We have argued with fictitious electron orbits (current loops), but we have ended up with exactly what happens to a diamagnetic material: If we apply the magnetic field of Fig. 32-12, the material develops a downward magnetic dipole moment and experiences an upward force. When the field is removed, both the dipole moment and the force disappear. The external field need not be positioned as shown in Fig. 32-12; similar arguments can be made for other orientations of \vec{B}_{ext} . In general,



A diamagnetic material placed in an external magnetic field \vec{B}_{ext} develops a magnetic dipole moment directed opposite \vec{B}_{ext} . If the field is nonuniform, the diamagnetic material is repelled *from* a region of greater magnetic field *toward* a region of lesser field.

The frog in Fig. 32-13 is diamagnetic (as is any other animal). When the frog was placed in the diverging magnetic field near the top end of a vertical current-carrying solenoid, every atom in the frog was repelled upward, away from the region of stronger magnetic field at that end of the solenoid. The frog moved upward into weaker and weaker magnetic field until the upward magnetic force balanced the gravitational force on it, and there it hung in midair. The frog is not in discomfort because *every* atom is subject to the same forces and thus there is no force variation within the frog. The sensation is similar to the “weightless” situation of floating in water, which frogs like very much. If we went to the expense of building a much larger solenoid, we could similarly levitate a person in midair due to the person’s diamagnetism.



CHECKPOINT 5

The figure shows two diamagnetic spheres located near the south pole of a bar magnet. Are (a) the magnetic forces on the spheres and (b) the magnetic dipole moments of the spheres directed toward or away from the bar magnet? (c) Is the magnetic force on sphere 1 greater than, less than, or equal to that on sphere 2?



32-10 Paramagnetism

In paramagnetic materials, the spin and orbital magnetic dipole moments of the electrons in each atom do not cancel but add vectorially to give the atom a net (and permanent) magnetic dipole moment $\vec{\mu}$. In the absence of an external magnetic field, these atomic dipole moments are randomly oriented, and the net magnetic dipole moment of the material is zero. However, if a sample of the material is placed in an external magnetic field \vec{B}_{ext} , the magnetic dipole moments tend to line up with the field, which gives the sample a net magnetic dipole moment. This alignment with the external field is the opposite of what we saw with diamagnetic materials.



A paramagnetic material placed in an external magnetic field \vec{B}_{ext} develops a magnetic dipole moment in the direction of \vec{B}_{ext} . If the field is nonuniform, the paramagnetic material is attracted *toward* a region of greater magnetic field *from* a region of lesser field.

A paramagnetic sample with N atoms would have a magnetic dipole moment of magnitude $N\mu$ if alignment of its atomic dipoles were complete. However, random collisions of atoms due to their thermal agitation transfer energy among the atoms, disrupting their alignment and thus reducing the sample's magnetic dipole moment.

The importance of thermal agitation may be measured by comparing two energies. One, given by Eq. 19-24, is the mean translational kinetic energy $K (= \frac{3}{2}kT)$ of an atom at temperature T , where k is the Boltzmann constant (1.38×10^{-23} J/K) and T is in kelvins (not Celsius degrees). The other, derived from Eq. 28-38, is the difference in energy $\Delta U_B (= 2\mu B_{\text{ext}})$ between parallel alignment and antiparallel alignment of the magnetic dipole moment of an atom and the external field. (The lower energy state is $-\mu B_{\text{ext}}$ and the higher energy state is $+\mu B_{\text{ext}}$.) As we shall show below, $K \gg \Delta U_B$, even for ordinary temperatures and field magnitudes. Thus, energy transfers during collisions among atoms can significantly disrupt the alignment of the atomic dipole moments, keeping the magnetic dipole moment of a sample much less than $N\mu$.

We can express the extent to which a given paramagnetic sample is magnetized by finding the ratio of its magnetic dipole moment to its volume V . This vector quantity, the magnetic dipole moment per unit volume, is the **magnetization** \vec{M} of the sample, and its magnitude is

$$M = \frac{\text{measured magnetic moment}}{V}. \quad (32-38)$$

The unit of \vec{M} is the ampere-square meter per cubic meter, or ampere per meter (A/m). Complete alignment of the atomic dipole moments, called *saturation* of the sample, corresponds to the maximum value $M_{\text{max}} = N\mu/V$.

In 1895 Pierre Curie discovered experimentally that the magnetization of a paramagnetic sample is directly proportional to the magnitude of the external magnetic field B_{ext} and inversely proportional to the temperature T in kelvins:

$$M = C \frac{B_{\text{ext}}}{T}. \quad (32-39)$$

Equation 32-39 is known as *Curie's law*, and C is called the *Curie constant*. Curie's law is reasonable in that increasing B_{ext} tends to align the atomic dipole moments in a sample and thus to increase M , whereas increasing T tends to disrupt the alignment via thermal agitation and thus to decrease M . However, the law is actually an approximation that is valid only when the ratio B_{ext}/T is not too large.



Liquid oxygen is suspended between the two pole faces of a magnet because the liquid is paramagnetic and is magnetically attracted to the magnet. (Richard Megna/Fundamental Photographs)

Fig. 32-14 A magnetization curve for potassium chromium sulfate, a paramagnetic salt. The ratio of magnetization M of the salt to the maximum possible magnetization M_{\max} is plotted versus the ratio of the applied magnetic field magnitude B_{ext} to the temperature T . Curie's law fits the data at the left; quantum theory fits all the data. After W. E. Henry.

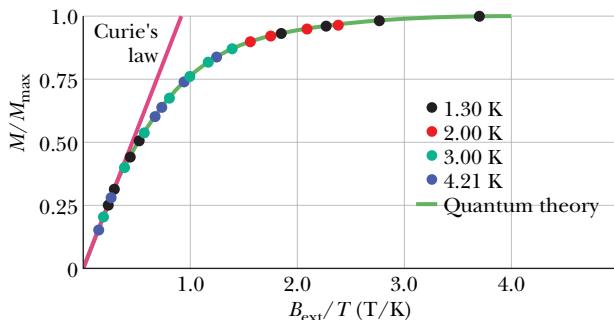


Figure 32-14 shows the ratio M/M_{\max} as a function of B_{ext}/T for a sample of the salt potassium chromium sulfate, in which chromium ions are the paramagnetic substance. The plot is called a *magnetization curve*. The straight line for Curie's law fits the experimental data at the left, for B_{ext}/T below about 0.5 T/K. The curve that fits all the data points is based on quantum physics. The data on the right side, near saturation, are very difficult to obtain because they require very strong magnetic fields (about 100 000 times Earth's field), even at very low temperatures.

CHECKPOINT 6

The figure here shows two paramagnetic spheres located near the south pole of a bar magnet. Are (a) the magnetic forces on the spheres and (b) the magnetic dipole moments of the spheres directed toward or away from the bar magnet? (c) Is the magnetic force on sphere 1 greater than, less than, or equal to that on sphere 2?



Sample Problem

Orientation energy of a magnetic field in a paramagnetic gas

A paramagnetic gas at room temperature ($T = 300 \text{ K}$) is placed in an external uniform magnetic field of magnitude $B = 1.5 \text{ T}$; the atoms of the gas have magnetic dipole moment $\mu = 1.0 \mu_B$. Calculate the mean translational kinetic energy K of an atom of the gas and the energy difference ΔU_B between parallel alignment and antiparallel alignment of the atom's magnetic dipole moment with the external field.

KEY IDEAS

- (1) The mean translational kinetic energy K of an atom in a gas depends on the temperature of the gas. (2) The energy U_B of a magnetic dipole $\vec{\mu}$ in an external magnetic field \vec{B} depends on the angle θ between the directions of $\vec{\mu}$ and \vec{B} .

Calculations: From Eq. 19-24, we have

$$K = \frac{3}{2}kT = \frac{3}{2}(1.38 \times 10^{-23} \text{ J/K})(300 \text{ K}) \\ = 6.2 \times 10^{-21} \text{ J} = 0.039 \text{ eV.} \quad (\text{Answer})$$

From Eq. 28-38 ($U_B = -\vec{\mu} \cdot \vec{B}$), we can write the difference ΔU_B between parallel alignment ($\theta = 0^\circ$) and antiparallel alignment ($\theta = 180^\circ$) as

$$\begin{aligned} \Delta U_B &= -\mu B \cos 180^\circ - (-\mu B \cos 0^\circ) = 2\mu B \\ &= 2\mu_B B = 2(9.27 \times 10^{-24} \text{ J/T})(1.5 \text{ T}) \\ &= 2.8 \times 10^{-23} \text{ J} = 0.00017 \text{ eV.} \quad (\text{Answer}) \end{aligned}$$

Here K is about 230 times ΔU_B ; so energy exchanges among the atoms during their collisions with one another can easily reorient any magnetic dipole moments that might be aligned with the external magnetic field. That is, as soon as a magnetic dipole moment happens to become aligned with the external field, in the dipole's low energy state, chances are very good that a neighboring atom will hit the atom, transferring enough energy to put the dipole in a higher energy state. Thus, the magnetic dipole moment exhibited by the paramagnetic gas must be due to fleeting partial alignments of the atomic dipole moments.

32-11 Ferromagnetism

When we speak of magnetism in everyday conversation, we almost always have a mental picture of a bar magnet or a disk magnet (probably clinging to a refrigerator door). That is, we picture a ferromagnetic material having strong, permanent magnetism, and not a diamagnetic or paramagnetic material having weak, temporary magnetism.

Iron, cobalt, nickel, gadolinium, dysprosium, and alloys containing these elements exhibit ferromagnetism because of a quantum physical effect called *exchange coupling* in which the electron spins of one atom interact with those of neighboring atoms. The result is alignment of the magnetic dipole moments of the atoms, in spite of the randomizing tendency of atomic collisions due to thermal agitation. This persistent alignment is what gives ferromagnetic materials their permanent magnetism.

If the temperature of a ferromagnetic material is raised above a certain critical value, called the *Curie temperature*, the exchange coupling ceases to be effective. Most such materials then become simply paramagnetic; that is, the dipoles still tend to align with an external field but much more weakly, and thermal agitation can now more easily disrupt the alignment. The Curie temperature for iron is 1043 K (= 770°C).

The magnetization of a ferromagnetic material such as iron can be studied with an arrangement called a *Rowland ring* (Fig. 32-15). The material is formed into a thin toroidal core of circular cross section. A primary coil P having n turns per unit length is wrapped around the core and carries current i_P . (The coil is essentially a long solenoid bent into a circle.) If the iron core were not present, the magnitude of the magnetic field inside the coil would be, from Eq. 29-23,

$$B_0 = \mu_0 i_P n. \quad (32-40)$$

However, with the iron core present, the magnetic field \vec{B} inside the coil is greater than \vec{B}_0 , usually by a large amount. We can write the magnitude of this field as

$$B = B_0 + B_M, \quad (32-41)$$

where B_M is the magnitude of the magnetic field contributed by the iron core. This contribution results from the alignment of the atomic dipole moments within the iron, due to exchange coupling and to the applied magnetic field B_0 , and is proportional to the magnetization M of the iron. That is, the contribution B_M is proportional to the magnetic dipole moment per unit volume of the iron. To determine B_M we use a secondary coil S to measure B , compute B_0 with Eq. 32-40, and subtract as suggested by Eq. 32-41.

Figure 32-16 shows a magnetization curve for a ferromagnetic material in a Rowland ring: The ratio $B_M/B_{M,\max}$, where $B_{M,\max}$ is the maximum possible value of B_M , corresponding to saturation, is plotted versus B_0 . The curve is like Fig. 32-14, the magnetization curve for a paramagnetic substance: Both curves show the extent to which an applied magnetic field can align the atomic dipole moments of a material.

For the ferromagnetic core yielding Fig. 32-16, the alignment of the dipole moments is about 70% complete for $B_0 \approx 1 \times 10^{-3}$ T. If B_0 were increased to 1 T, the alignment would be almost complete (but $B_0 = 1$ T, and thus almost complete saturation, is quite difficult to obtain).

Magnetic Domains

Exchange coupling produces strong alignment of adjacent atomic dipoles in a ferromagnetic material at a temperature below the Curie temperature. Why, then, isn't the material naturally at saturation even when there is no applied magnetic field B_0 ? Why isn't every piece of iron a naturally strong magnet?

To understand this, consider a specimen of a ferromagnetic material such as iron that is in the form of a single crystal; that is, the arrangement of the atoms that

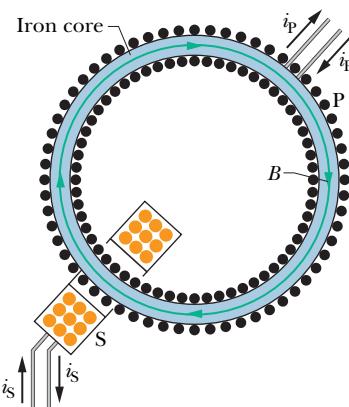


Fig. 32-15 A Rowland ring. A primary coil P has a core made of the ferromagnetic material to be studied (here iron). The core is magnetized by a current i_P sent through coil P. (The turns of the coil are represented by dots.) The extent to which the core is magnetized determines the total magnetic field \vec{B} within coil P. Field \vec{B} can be measured by means of a secondary coil S.

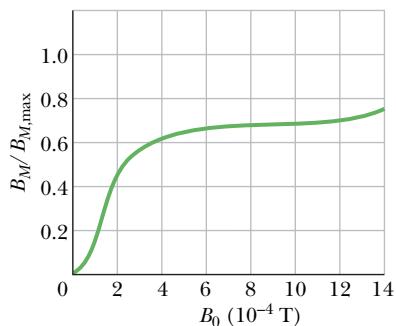


Fig. 32-16 A magnetization curve for a ferromagnetic core material in the Rowland ring of Fig. 32-15. On the vertical axis, 1.0 corresponds to complete alignment (saturation) of the atomic dipoles within the material.

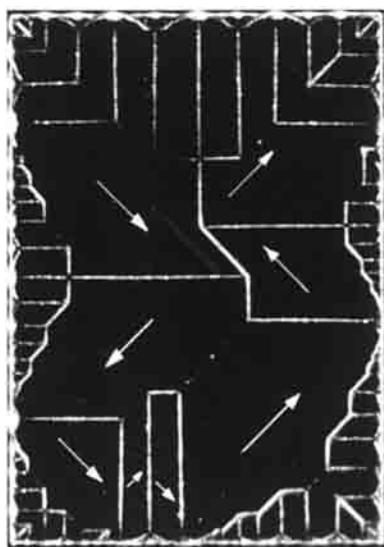


Fig. 32-17 A photograph of domain patterns within a single crystal of nickel; white lines reveal the boundaries of the domains. The white arrows superimposed on the photograph show the orientations of the magnetic dipoles within the domains and thus the orientations of the net magnetic dipoles of the domains. The crystal as a whole is unmagnetized if the net magnetic field (the vector sum over all the domains) is zero. (Courtesy Ralph W. DeBlois)

make it up—its crystal lattice—extends with unbroken regularity throughout the volume of the specimen. Such a crystal will, in its normal state, be made up of a number of *magnetic domains*. These are regions of the crystal throughout which the alignment of the atomic dipoles is essentially perfect. The domains, however, are not all aligned. For the crystal as a whole, the domains are so oriented that they largely cancel with one another as far as their external magnetic effects are concerned.

Figure 32-17 is a magnified photograph of such an assembly of domains in a single crystal of nickel. It was made by sprinkling a colloidal suspension of finely powdered iron oxide on the surface of the crystal. The domain boundaries, which are thin regions in which the alignment of the elementary dipoles changes from a certain orientation in one of the domains forming the boundary to a different orientation in the other domain, are the sites of intense, but highly localized and nonuniform, magnetic fields. The suspended colloidal particles are attracted to these boundaries and show up as the white lines (not all the domain boundaries are apparent in Fig. 32-17). Although the atomic dipoles in each domain are completely aligned as shown by the arrows, the crystal as a whole may have only a very small resultant magnetic moment.

Actually, a piece of iron as we ordinarily find it is not a single crystal but an assembly of many tiny crystals, randomly arranged; we call it a *polycrystalline solid*. Each tiny crystal, however, has its array of variously oriented domains, just as in Fig. 32-17. If we magnetize such a specimen by placing it in an external magnetic field of gradually increasing strength, we produce two effects; together they produce a magnetization curve of the shape shown in Fig. 32-16. One effect is a growth in size of the domains that are oriented along the external field at the expense of those that are not. The second effect is a shift of the orientation of the dipoles within a domain, as a unit, to become closer to the field direction.

Exchange coupling and domain shifting give us the following result:



A ferromagnetic material placed in an external magnetic field \vec{B}_{ext} develops a strong magnetic dipole moment in the direction of \vec{B}_{ext} . If the field is nonuniform, the ferromagnetic material is attracted *toward* a region of greater magnetic field *from* a region of lesser field.

Hysteresis

Magnetization curves for ferromagnetic materials are not retraced as we increase and then decrease the external magnetic field B_0 . Figure 32-18 is a plot of B_M versus B_0 during the following operations with a Rowland ring: (1) Starting with the iron unmagnetized (point *a*), increase the current in the toroid until B_0 ($= \mu_0 i n$) has the value corresponding to point *b*; (2) reduce the current in the toroid winding (and thus B_0) back to zero (point *c*); (3) reverse the toroid current and increase it in magnitude until B_0 has the value corresponding to point *d*; (4) reduce the current to zero again (point *e*); (5) reverse the current once more until point *b* is reached again.

The lack of retraceability shown in Fig. 32-18 is called **hysteresis**, and the curve *bcdeb* is called a *hysteresis loop*. Note that at points *c* and *e* the iron core is magnetized, even though there is no current in the toroid windings; this is the familiar phenomenon of permanent magnetism.

Hysteresis can be understood through the concept of magnetic domains. Evidently the motions of the domain boundaries and the reorientations of the domain directions are not totally reversible. When the applied magnetic field B_0 is increased and then decreased back to its initial value, the domains do not return completely to their original configuration but retain some “memory” of their alignment after the initial increase. This memory of magnetic materials is essential for the magnetic storage of information.

This memory of the alignment of domains can also occur naturally. When lightning sends currents along multiple tortuous paths through the ground,

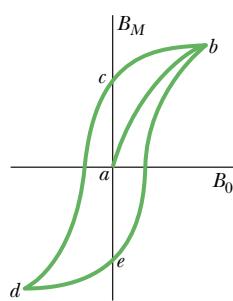


Fig. 32-18 A magnetization curve (*ab*) for a ferromagnetic specimen and an associated hysteresis loop (*bcdeb*).

the currents produce intense magnetic fields that can suddenly magnetize any ferromagnetic material in nearby rock. Because of hysteresis, such rock material retains some of that magnetization after the lightning strike (after the currents disappear). Pieces of the rock—later exposed, broken, and loosened by weathering—are then lodestones.

Sample Problem

Magnetic dipole moment of a compass needle

A compass needle made of pure iron (density 7900 kg/m^3) has a length L of 3.0 cm, a width of 1.0 mm, and a thickness of 0.50 mm. The magnitude of the magnetic dipole moment of an iron atom is $\mu_{\text{Fe}} = 2.1 \times 10^{-23} \text{ J/T}$. If the magnetization of the needle is equivalent to the alignment of 10% of the atoms in the needle, what is the magnitude of the needle's magnetic dipole moment $\vec{\mu}$?

KEY IDEAS

(1) Alignment of all N atoms in the needle would give a magnitude of $N\mu_{\text{Fe}}$ for the needle's magnetic dipole moment $\vec{\mu}$. However, the needle has only 10% alignment (the random orientation of the rest does not give any net contribution to $\vec{\mu}$). Thus,

$$\mu = 0.10N\mu_{\text{Fe}}. \quad (32-42)$$

(2) We can find the number of atoms N in the needle from the needle's mass:

$$N = \frac{\text{needle's mass}}{\text{iron's atomic mass}}. \quad (32-43)$$

Finding N : Iron's atomic mass is not listed in Appendix F, but its molar mass M is. Thus, we write

$$\text{iron's atomic mass} = \frac{\text{iron's molar mass } M}{\text{Avogadro's number } N_A}. \quad (32-44)$$



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Next, we can rewrite Eq. 32-43 in terms of the needle's mass m , the molar mass M , and Avogadro's number N_A :

$$N = \frac{mN_A}{M}. \quad (32-45)$$

The needle's mass m is the product of its density and its volume. The volume works out to be $1.5 \times 10^{-8} \text{ m}^3$; so

$$\begin{aligned} \text{needle's mass } m &= (\text{needle's density})(\text{needle's volume}) \\ &= (7900 \text{ kg/m}^3)(1.5 \times 10^{-8} \text{ m}^3) \\ &= 1.185 \times 10^{-4} \text{ kg}. \end{aligned}$$

Substituting into Eq. 32-45 with this value for m , and also 55.847 g/mol ($= 0.055847 \text{ kg/mol}$) for M and 6.02×10^{23} for N_A , we find

$$\begin{aligned} N &= \frac{(1.185 \times 10^{-4} \text{ kg})(6.02 \times 10^{23})}{0.055847 \text{ kg/mol}} \\ &= 1.2774 \times 10^{21}. \end{aligned}$$

Finding μ : Substituting our value of N and the value of μ_{Fe} into Eq. 32-42 then yields

$$\begin{aligned} \mu &= (0.10)(1.2774 \times 10^{21})(2.1 \times 10^{-23} \text{ J/T}) \\ &= 2.682 \times 10^{-3} \text{ J/T} \approx 2.7 \times 10^{-3} \text{ J/T}. \quad (\text{Answer}) \end{aligned}$$

REVIEW & SUMMARY

Gauss' Law for Magnetic Fields The simplest magnetic structures are magnetic dipoles. Magnetic monopoles do not exist (as far as we know). **Gauss' law** for magnetic fields,

$$\oint \vec{B} \cdot d\vec{A} = 0, \quad (32-1)$$

states that the net magnetic flux through any (closed) Gaussian surface is zero. It implies that magnetic monopoles do not exist.

Maxwell's Extension of Ampere's Law A changing electric flux induces a magnetic field \vec{B} . Maxwell's law,

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt} \quad (\text{Maxwell's law of induction}), \quad (32-3)$$

relates the magnetic field induced along a closed loop to the changing electric flux Φ_E through the loop. Ampere's law, $\oint \vec{B} \cdot d\vec{s} = \mu_0 i_{\text{enc}}$ (Eq. 32-4), gives the magnetic field generated by a current i_{enc} encircled by a closed loop. Maxwell's law and Ampere's law can be written as the single equation

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt} + \mu_0 i_{\text{enc}} \quad (\text{Ampere-Maxwell law}). \quad (32-5)$$

Displacement Current We define the fictitious *displacement current* due to a changing electric field as

$$i_d = \epsilon_0 \frac{d\Phi_E}{dt}. \quad (32-10)$$

Equation 32-5 then becomes

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 i_{d,\text{enc}} + \mu_0 i_{\text{enc}} \quad (\text{Ampere-Maxwell law}), \quad (32-11)$$

where $i_{d,\text{enc}}$ is the displacement current encircled by the integration loop. The idea of a displacement current allows us to retain the notion of continuity of current through a capacitor. However, displacement current is *not* a transfer of charge.

Maxwell's Equations Maxwell's equations, displayed in Table 32-1, summarize electromagnetism and form its foundation, including optics.

Earth's Magnetic Field Earth's magnetic field can be approximated as being that of a magnetic dipole whose dipole moment makes an angle of 11.5° with Earth's rotation axis, and with the south pole of the dipole in the Northern Hemisphere. The direction of the local magnetic field at any point on Earth's surface is given by the *field declination* (the angle left or right from geographic north) and the *field inclination* (the angle up or down from the horizontal).

Spin Magnetic Dipole Moment An electron has an intrinsic angular momentum called *spin angular momentum* (or *spin*) \vec{S} , with which an intrinsic *spin magnetic dipole moment* $\vec{\mu}_s$ is associated:

$$\vec{\mu}_s = -\frac{e}{m} \vec{S}. \quad (32-22)$$

Spin \vec{S} cannot itself be measured, but any component can be measured. Assuming that the measurement is along the z axis of a coordinate system, the component S_z can have only the values given by

$$S_z = m_s \frac{h}{2\pi}, \quad \text{for } m_s = \pm \frac{1}{2}, \quad (32-23)$$

where $h (= 6.63 \times 10^{-34} \text{ J}\cdot\text{s})$ is the Planck constant. Similarly, the electron's spin magnetic dipole moment $\vec{\mu}_s$ cannot itself be measured but its component can be measured. Along a z axis, the component is

$$\mu_{s,z} = \pm \frac{eh}{4\pi m} = \pm \mu_B, \quad (32-24, 32-26)$$

where μ_B is the *Bohr magneton*:

$$\mu_B = \frac{eh}{4\pi m} = 9.27 \times 10^{-24} \text{ J/T}. \quad (32-25)$$

The energy U associated with the orientation of the spin magnetic dipole moment in an external magnetic field \vec{B}_{ext} is

$$U = -\vec{\mu}_s \cdot \vec{B}_{\text{ext}} = -\mu_{s,z} B_{\text{ext}}. \quad (32-27)$$

Orbital Magnetic Dipole Moment An electron in an atom has an additional angular momentum called its *orbital angular momentum* \vec{L}_{orb} , with which an *orbital magnetic dipole moment* $\vec{\mu}_{\text{orb}}$ is associated:

$$\vec{\mu}_{\text{orb}} = -\frac{e}{2m} \vec{L}_{\text{orb}}. \quad (32-28)$$

Orbital angular momentum is quantized and can have only values

given by

$$L_{\text{orb},z} = m_\ell \frac{h}{2\pi},$$

for $m_\ell = 0, \pm 1, \pm 2, \dots, \pm (\text{limit}).$ (32-29)

This means that the associated magnetic dipole moment measured along a z axis is given by

$$\mu_{\text{orb},z} = -m_\ell \frac{eh}{4\pi m} = -m_\ell \mu_B. \quad (32-30, 32-31)$$

The energy U associated with the orientation of the orbital magnetic dipole moment in an external magnetic field \vec{B}_{ext} is

$$U = -\vec{\mu}_{\text{orb}} \cdot \vec{B}_{\text{ext}} = -\mu_{\text{orb},z} B_{\text{ext}}. \quad (32-32)$$

Diamagnetism *Diamagnetic materials* do not exhibit magnetism until they are placed in an external magnetic field \vec{B}_{ext} . They then develop a magnetic dipole moment directed opposite \vec{B}_{ext} . If the field is nonuniform, the diamagnetic material is repelled from regions of greater magnetic field. This property is called *diamagnetism*.

Paramagnetism In a *paramagnetic material*, each atom has a permanent magnetic dipole moment $\vec{\mu}$, but the dipole moments are randomly oriented and the material as a whole lacks a magnetic field. However, an external magnetic field \vec{B}_{ext} can partially align the atomic dipole moments to give the material a net magnetic dipole moment in the direction of \vec{B}_{ext} . If \vec{B}_{ext} is nonuniform, the material is attracted to regions of greater magnetic field. These properties are called *paramagnetism*.

The alignment of the atomic dipole moments increases with an increase in \vec{B}_{ext} and decreases with an increase in temperature T . The extent to which a sample of volume V is magnetized is given by its *magnetization* \vec{M} , whose magnitude is

$$M = \frac{\text{measured magnetic moment}}{V}. \quad (32-38)$$

Complete alignment of all N atomic magnetic dipoles in a sample, called *saturation* of the sample, corresponds to the maximum magnetization value $M_{\text{max}} = N\mu/V$. For low values of the ratio B_{ext}/T , we have the approximation

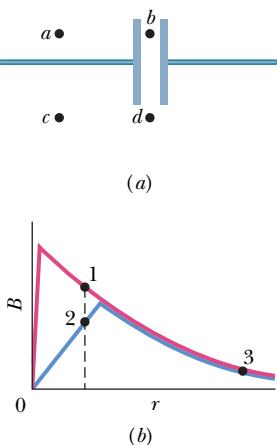
$$M = C \frac{B_{\text{ext}}}{T} \quad (\text{Curie's law}), \quad (32-39)$$

where C is called the *Curie constant*.

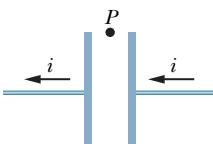
Ferromagnetism In the absence of an external magnetic field, some of the electrons in a ferromagnetic material have their magnetic dipole moments aligned by means of a quantum physical interaction called *exchange coupling*, producing regions (domains) within the material with strong magnetic dipole moments. An external field \vec{B}_{ext} can align the magnetic dipole moments of those regions, producing a strong net magnetic dipole moment for the material as a whole, in the direction of \vec{B}_{ext} . This net magnetic dipole moment can partially persist when field \vec{B}_{ext} is removed. If \vec{B}_{ext} is nonuniform, the ferromagnetic material is attracted to regions of greater magnetic field. These properties are called *ferromagnetism*. Exchange coupling disappears when a sample's temperature exceeds its *Curie temperature*.

QUESTIONS

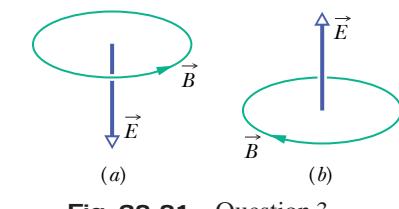
- 1** Figure 32-19a shows a capacitor, with circular plates, that is being charged. Point *a* (near one of the connecting wires) and point *b* (inside the capacitor gap) are equidistant from the central axis, as are point *c* (not so near the wire) and point *d* (between the plates but outside the gap). In Fig. 32-19b, one curve gives the variation with distance *r* of the magnitude of the magnetic field inside and outside the wire. The other curve gives the variation with distance *r* of the magnitude of the magnetic field inside and outside the gap. The two curves partially overlap. Which of the three points on the curves correspond to which of the four points of Fig. 32-19a?

**Fig. 32-19** Question 1.

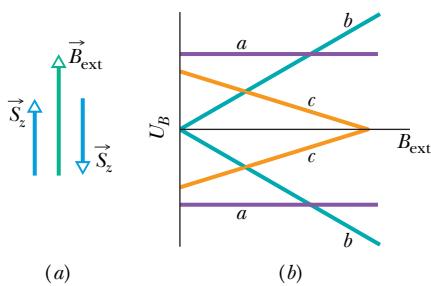
- 2** Figure 32-20 shows a parallel-plate capacitor and the current in the connecting wires that is discharging the capacitor. Are the directions of (a) electric field \vec{E} and (b) displacement current i_d leftward or rightward between the plates? (c) Is the magnetic field at point *P* into or out of the page?

**Fig. 32-20** Question 2.

- 3** Figure 32-21 shows, in two situations, an electric field vector \vec{E} and an induced magnetic field line. In each, is the magnitude of \vec{E} increasing or decreasing?

**Fig. 32-21** Question 3.

- 4** Figure 32-22a shows a pair of opposite spin orientations for an electron in an external magnetic field \vec{B}_{ext} . Figure 32-22b gives three choices for the graph of the potential energies associated with those orientations as a function of the magnitude of \vec{B}_{ext} . Choices *b* and *c* consist of intersecting lines, choice *a* of parallel lines. Which is the correct choice?

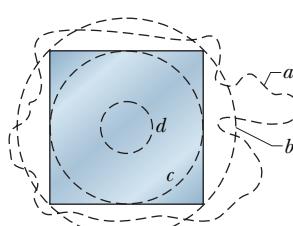
**Fig. 32-22** Question 4.

- 5** An electron in an external magnetic field \vec{B}_{ext} has its spin angular momentum S_z antiparallel to \vec{B}_{ext} . If the electron undergoes a *spin-flip* so that S_z is then parallel with \vec{B}_{ext} , must energy be supplied to or lost by the electron?

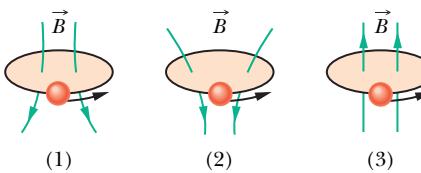
- 6** Does the magnitude of the net force on the current loop of Figs. 32-12a and b increase, decrease, or remain the same if we increase (a) the magnitude of \vec{B}_{ext} and (b) the divergence of \vec{B}_{ext} ?

- 7** Figure 32-23 shows a face-on view of one of the two square plates of a parallel-plate capacitor, as well as four loops that are located between the plates. The capacitor is being discharged. (a)

- Neglecting fringing of the magnetic field, rank the loops according to the magnitude of $\oint \vec{B} \cdot d\vec{s}$ along them, greatest first. (b) Along which loop, if any, is the angle between the directions of \vec{B} and $d\vec{s}$ constant (so that their dot product can easily be evaluated)? (c) Along which loop, if any, is B constant (so that B can be brought in front of the integral sign in Eq. 32-3)?

**Fig. 32-23** Question 7.

- 8** Figure 32-24 shows three loop models of an electron orbiting counterclockwise within a magnetic field. The fields are nonuniform for models 1 and 2 and uniform for model 3. For each model, are (a) the magnetic dipole moment of the loop and (b) the magnetic force on the loop directed up, directed down, or zero?

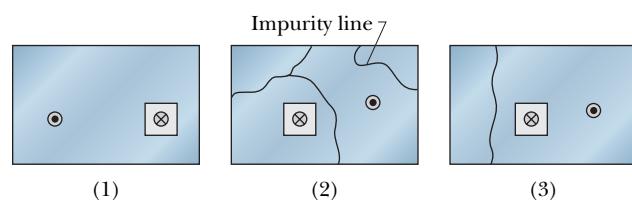
**Fig. 32-24** Questions 8, 9, and 10.

- 9** Replace the current loops of Question 8 and Fig. 32-24 with diamagnetic spheres. For each field, are (a) the magnetic dipole moment of the sphere and (b) the magnetic force on the sphere directed up, directed down, or zero?

- 10** Replace the current loops of Question 8 and Fig. 32-24 with paramagnetic spheres. For each field, are (a) the magnetic dipole moment of the sphere and (b) the magnetic force on the sphere directed up, directed down, or zero?

- 11** Figure 32-25 represents three rectangular samples of a ferromagnetic material in which the magnetic dipoles of the domains have been directed out of the page (encircled dot) by a very strong applied field B_0 . In each sample, an island domain still has its magnetic field directed into the page (encircled \times). Sample 1 is one (pure) crystal. The other samples contain impurities collected along lines; domains cannot easily spread across such lines.

- The applied field is now to be reversed and its magnitude kept moderate. The change causes the island domain to grow. (a) Rank the three samples according to the success of that growth, greatest growth first. Ferromagnetic materials in which the magnetic dipoles are easily changed are said to be *magnetically soft*; when the changes are difficult, requiring strong applied fields, the materials are said to be *magnetically hard*. (b) Of the three samples, which is the most magnetically hard?

**Fig. 32-25** Question 11.

PROBLEMS



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com



sec. 32-2 Gauss' Law for Magnetic Fields

- 1 The magnetic flux through each of five faces of a die (singular of “dice”) is given by $\Phi_B = \pm N \text{ WB}$, where N ($= 1$ to 5) is the number of spots on the face. The flux is positive (outward) for N even and negative (inward) for N odd. What is the flux through the sixth face of the die?

- 2 Figure 32-26 shows a closed surface. Along the flat top face, which has a radius of 2.0 cm , a perpendicular magnetic field \vec{B} of magnitude 0.30 T is directed outward. Along the flat bottom face, a magnetic flux of 0.70 mWB is directed outward. What are the (a) magnitude and (b) direction (inward or outward) of the magnetic flux through the curved part of the surface?

- 3 **SSM ILW** A Gaussian surface in the shape of a right circular cylinder with end caps has a radius of 12.0 cm and a length of 80.0 cm . Through one end there is an inward magnetic flux of $25.0 \mu\text{WB}$. At the other end there is a uniform magnetic field of 1.60 mT , normal to the surface and directed outward. What are the (a) magnitude and (b) direction (inward or outward) of the net magnetic flux through the curved surface?

- 4 Two wires, parallel to a z axis and a distance $4r$ apart, carry equal currents i in opposite directions, as shown in Fig. 32-27. A circular cylinder of radius r and length L has its axis on the z axis, midway between the wires. Use Gauss' law for magnetism to derive an expression for the net outward magnetic flux through the half of the cylindrical surface above the x axis. (Hint: Find the flux through the portion of the xz plane that lies within the cylinder.)

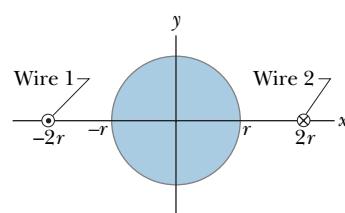


Fig. 32-27 Problem 4.

sec. 32-3 Induced Magnetic Fields

- 5 **SSM** The induced magnetic field at radial distance 6.0 mm from the central axis of a circular parallel-plate capacitor is $2.0 \times 10^{-7} \text{ T}$. The plates have radius 3.0 mm . At what rate dE/dt is the electric field between the plates changing?

- 6 A capacitor with square plates of edge length L is being discharged by a current of 0.75 A . Figure 32-28 is a head-on view of one of the plates from inside the capacitor. A dashed rectangular path is shown. If $L = 12 \text{ cm}$, $W = 4.0 \text{ cm}$, and $H = 2.0 \text{ cm}$, what is the value of $\oint \vec{B} \cdot d\vec{s}$ around the dashed path?

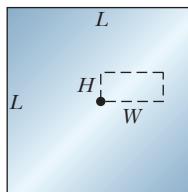


Fig. 32-28
Problem 6.

- 7 **GO** Uniform electric flux. Figure 32-29 shows a circular region of radius $R = 3.00 \text{ cm}$ in which a uniform electric flux is directed out of the plane of the

page. The total electric flux through the region is given by $\Phi_E = (3.00 \text{ mV} \cdot \text{m/s})t$, where t is in seconds. What is the magnitude of the magnetic field that is induced at radial distances (a) 2.00 cm and (b) 5.00 cm ?

- 8 **GO** Nonuniform electric flux. Figure 32-29 shows a circular region of radius $R = 3.00 \text{ cm}$ in which an electric flux is directed out of the plane of the page. The flux encircled by a concentric circle of radius r is given by $\Phi_{E,\text{enc}} = (0.600 \text{ V} \cdot \text{m/s})(r/R)t$, where $r \leq R$ and t is in seconds. What is the magnitude of the induced magnetic field at radial distances (a) 2.00 cm and (b) 5.00 cm ?

- 9 **GO** Uniform electric field. In Fig. 32-29, a uniform electric field is directed out of the page within a circular region of radius $R = 3.00 \text{ cm}$. The field magnitude is given by $E = (4.50 \times 10^{-3} \text{ V/m} \cdot \text{s})t$, where t is in seconds. What is the magnitude of the induced magnetic field at radial distances (a) 2.00 cm and (b) 5.00 cm ?

- 10 **GO** Nonuniform electric field. In Fig. 32-29, an electric field is directed out of the page within a circular region of radius $R = 3.00 \text{ cm}$. The field magnitude is $E = (0.500 \text{ V/m} \cdot \text{s})(1 - r/R)t$, where t is in seconds and r is the radial distance ($r \leq R$). What is the magnitude of the induced magnetic field at radial distances (a) 2.00 cm and (b) 5.00 cm ?

- 11 Suppose that a parallel-plate capacitor has circular plates with radius $R = 30 \text{ mm}$ and a plate separation of 5.0 mm . Suppose also that a sinusoidal potential difference with a maximum value of 150 V and a frequency of 60 Hz is applied across the plates; that is,

$$V = (150 \text{ V}) \sin[2\pi(60 \text{ Hz})t].$$

- (a) Find $B_{\max}(R)$, the maximum value of the induced magnetic field that occurs at $r = R$. (b) Plot $B_{\max}(r)$ for $0 < r < 10 \text{ cm}$.

- 12 A parallel-plate capacitor with circular plates of radius 40 mm is being discharged by a current of 6.0 A . At what radius (a) inside and (b) outside the capacitor gap is the magnitude of the induced magnetic field equal to 75% of its maximum value? (c) What is that maximum value?

sec. 32-4 Displacement Current

- 13 At what rate must the potential difference between the plates of a parallel-plate capacitor with a $2.0 \mu\text{F}$ capacitance be changed to produce a displacement current of 1.5 A ?

- 14 A parallel-plate capacitor with circular plates of radius R is being charged. Show that the magnitude of the current density of the displacement current is $J_d = \epsilon_0(dE/dt)$ for $r \leq R$.

- 15 **SSM** Prove that the displacement current in a parallel-plate capacitor of capacitance C can be written as $i_d = C(dV/dt)$, where V is the potential difference between the plates.

- 16 A parallel-plate capacitor with circular plates of radius 0.10 m is being discharged. A circular loop of radius 0.20 m is concentric

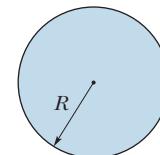


Fig. 32-29
Problems 7 to 10 and 19 to 22.

PROBLEMS

885

with the capacitor and halfway between the plates. The displacement current through the loop is 2.0 A. At what rate is the electric field between the plates changing?

••17 A silver wire has resistivity $\rho = 1.62 \times 10^{-8} \Omega \cdot \text{m}$ and a cross-sectional area of 5.00 mm^2 . The current in the wire is uniform and changing at the rate of 2000 A/s when the current is 100 A. (a) What is the magnitude of the (uniform) electric field in the wire when the current in the wire is 100 A? (b) What is the displacement current in the wire at that time? (c) What is the ratio of the magnitude of the magnetic field due to the displacement current to that due to the current at a distance r from the wire?

••18 The circuit in Fig. 32-30 consists of switch S, a 12.0 V ideal battery, a $20.0 \text{ M}\Omega$ resistor, and an air-filled capacitor. The capacitor has parallel circular plates of radius 5.00 cm , separated by 3.00 mm . At time $t = 0$, switch S is closed to begin charging the capacitor. The electric field between the plates is uniform. At $t = 250 \mu\text{s}$, what is the magnitude of the magnetic field within the capacitor, at radial distance 3.00 cm ?

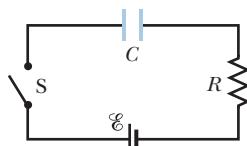


Fig. 32-30 Problem 18.

••19 Uniform displacement-current density. Figure 32-29 shows a circular region of radius $R = 3.00 \text{ cm}$ in which a displacement current is directed out of the page. The displacement current has a uniform density of magnitude $J_d = 6.00 \text{ A/m}^2$. What is the magnitude of the magnetic field due to the displacement current at radial distances (a) 2.00 cm and (b) 5.00 cm ?

••20 Uniform displacement current. Figure 32-29 shows a circular region of radius $R = 3.00 \text{ cm}$ in which a uniform displacement current $i_d = 0.500 \text{ A}$ is out of the page. What is the magnitude of the magnetic field due to the displacement current at radial distances (a) 2.00 cm and (b) 5.00 cm ?

••21 GO Nonuniform displacement-current density. Figure 32-29 shows a circular region of radius $R = 3.00 \text{ cm}$ in which a displacement current is directed out of the page. The magnitude of the density of this displacement current is $J_d = (4.00 \text{ A/m}^2)(1 - r/R)$, where r is the radial distance ($r \leq R$). What is the magnitude of the magnetic field due to the displacement current at (a) $r = 2.00 \text{ cm}$ and (b) $r = 5.00 \text{ cm}$?

••22 GO Nonuniform displacement current. Figure 32-29 shows a circular region of radius $R = 3.00 \text{ cm}$ in which a displacement current i_d is directed out of the page. The magnitude of the displacement current is given by $i_d = (3.00 \text{ A})(r/R)$, where r is the radial distance ($r \leq R$). What is the magnitude of the magnetic field due to i_d at radial distances (a) 2.00 cm and (b) 5.00 cm ?

••23 SSM ILW In Fig. 32-31, a parallel-plate capacitor has square plates of edge length $L = 1.0 \text{ m}$. A current of 2.0 A charges the capacitor, producing a uniform electric field \vec{E} between the plates, with \vec{E} perpendicular to the plates. (a) What is the displacement current i_d through the region between the plates? (b) What is dE/dt in this region? (c) What is the displacement current encircled by the square dashed path of edge length $d = 0.50 \text{ m}$? (d) What is $\oint \vec{B} \cdot d\vec{s}$ around this square dashed path?

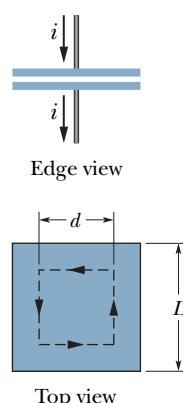


Fig. 32-31
Problem 23.

••24 The magnitude of the electric field between the two circular parallel plates in Fig. 32-32 is $E = (4.0 \times 10^5) - (6.0 \times 10^4 t)$, with E in volts per meter and t in seconds. At $t = 0$, \vec{E} is upward. The plate area is $4.0 \times 10^{-2} \text{ m}^2$. For $t \geq 0$, what are the (a) magnitude and (b) direction (up or down) of the displacement current between the plates and (c) is the direction of the induced magnetic field clockwise or counterclockwise in the figure?

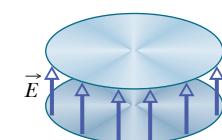


Fig. 32-32
Problem 24.

••25 ILW As a parallel-plate capacitor with circular plates 20 cm in diameter is being charged, the current density of the displacement current in the region between the plates is uniform and has a magnitude of 20 A/m^2 . (a) Calculate the magnitude B of the magnetic field at a distance $r = 50 \text{ mm}$ from the axis of symmetry of this region. (b) Calculate dE/dt in this region.

••26 A capacitor with parallel circular plates of radius $R = 1.20 \text{ cm}$ is discharging via a current of 12.0 A . Consider a loop of radius $R/3$ that is centered on the central axis between the plates. (a) How much displacement current is encircled by the loop? The maximum induced magnetic field has a magnitude of 12.0 mT . At what radius (b) inside and (c) outside the capacitor gap is the magnitude of the induced magnetic field 3.00 mT ?

••27 ILW In Fig. 32-33, a uniform electric field \vec{E} collapses. The vertical axis scale is set by $E_s = 6.0 \times 10^5 \text{ N/C}$, and the horizontal axis scale is set by $t_s = 12.0 \mu\text{s}$. Calculate the magnitude of the displacement current through a 1.6 m^2 area perpendicular to the field during each of the time intervals a , b , and c shown on the graph. (Ignore the behavior at the ends of the intervals.)

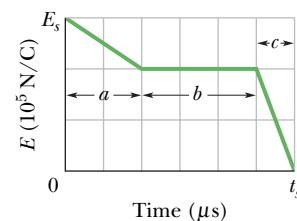


Fig. 32-33 Problem 27.

••28 GO Figure 32-34a shows the current i that is produced in a wire of resistivity $1.62 \times 10^{-8} \Omega \cdot \text{m}$. The magnitude of the current versus time t is shown in Fig. 32-34b. The vertical axis scale is set by $i_s = 10.0 \text{ A}$, and the horizontal axis scale is set by $t_s = 50.0 \text{ ms}$. Point P is at radial distance 9.00 mm from the wire's center. Determine the magnitude of the magnetic field \vec{B}_i at point P due to the actual current i in the wire at (a) $t = 20 \text{ ms}$, (b) $t = 40 \text{ ms}$, and (c) $t = 60 \text{ ms}$. Next, assume that the electric field driving the current is confined to the wire. Then determine the magnitude of the magnetic field \vec{B}_{id} at point P due to the displacement current i_d in the wire at (d) $t = 20 \text{ ms}$, (e) $t = 40 \text{ ms}$, and (f) $t = 60 \text{ ms}$. At point P at $t = 20 \text{ s}$, what is the direction (into or out of the page) of (g) \vec{B}_i and (h) \vec{B}_{id} ?

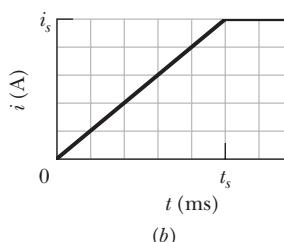
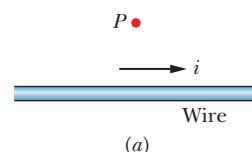


Fig. 32-34 Problem 28.

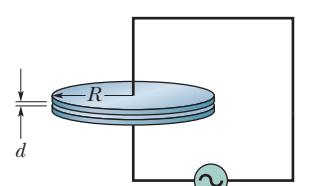


Fig. 32-35 Problem 29.

18.0 cm is connected to a source of emf $\mathcal{E} = \mathcal{E}_m \sin \omega t$, where $\mathcal{E}_m = 220$ V and $\omega = 130$ rad/s. The maximum value of the displacement current is $i_d = 7.60 \mu\text{A}$. Neglect fringing of the electric field at the edges of the plates. (a) What is the maximum value of the current i in the circuit? (b) What is the maximum value of $d\Phi_E/dt$, where Φ_E is the electric flux through the region between the plates? (c) What is the separation d between the plates? (d) Find the maximum value of the magnitude of \vec{B} between the plates at a distance $r = 11.0$ cm from the center.

sec. 32-6 Magnets

•30 Assume the average value of the vertical component of Earth's magnetic field is $43 \mu\text{T}$ (downward) for all of Arizona, which has an area of $2.95 \times 10^5 \text{ km}^2$. What then are the (a) magnitude and (b) direction (inward or outward) of the net magnetic flux through the rest of Earth's surface (the entire surface excluding Arizona)?

•31 In New Hampshire the average horizontal component of Earth's magnetic field in 1912 was $16 \mu\text{T}$, and the average inclination or "dip" was 73° . What was the corresponding magnitude of Earth's magnetic field?

sec. 32-7 Magnetism and Electrons

•32 Figure 32-36a is a one-axis graph along which two of the allowed energy values (*levels*) of an atom are plotted. When the atom is placed in a magnetic field of 0.500 T , the graph changes to that of Fig. 32-36b because of the energy associated with $\vec{\mu}_{\text{orb}} \cdot \vec{B}$. (We neglect $\vec{\mu}_s$.) Level E_1 is unchanged, but level E_2 splits into a (closely spaced) triplet of levels. What are the allowed values of m_ℓ associated with (a) energy level E_1 and (b) energy level E_2 ? (c) In joules, what amount of energy is represented by the spacing between the triplet levels?

•33 SSM WWW If an electron in an atom has an orbital angular momentum with $m = 0$, what are the components (a) $L_{\text{orb},z}$ and (b) $\mu_{\text{orb},z}$? If the atom is in an external magnetic field \vec{B} that has magnitude 35 mT and is directed along the z axis, what are (c) the energy U_{orb} associated with $\vec{\mu}_{\text{orb}}$ and (d) the energy U_{spin} associated with $\vec{\mu}_s$? If, instead, the electron has $m = -3$, what are (e) $L_{\text{orb},z}$, (f) $\mu_{\text{orb},z}$, (g) U_{orb} , and (h) U_{spin} ?

•34 What is the energy difference between parallel and antiparallel alignment of the z component of an electron's spin magnetic dipole moment with an external magnetic field of magnitude 0.25 T , directed parallel to the z axis?

•35 What is the measured component of the orbital magnetic dipole moment of an electron with (a) $m_\ell = 1$ and (b) $m_\ell = -2$?

•36 An electron is placed in a magnetic field \vec{B} that is directed along a z axis. The energy difference between parallel and antiparallel alignments of the z component of the electron's spin magnetic moment with \vec{B} is $6.00 \times 10^{-25} \text{ J}$. What is the magnitude of \vec{B} ?

sec. 32-9 Diamagnetism

•37 Figure 32-37 shows a loop model (loop L) for a diamagnetic material. (a) Sketch the magnetic field lines within and about the material due to the bar magnet. What is

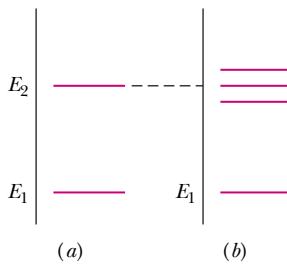


Fig. 32-36 Problem 32.

the direction of (b) the loop's net magnetic dipole moment $\vec{\mu}$, (c) the conventional current i in the loop (clockwise or counterclockwise in the figure), and (d) the magnetic force on the loop?

•38 Assume that an electron of mass m and charge magnitude e moves in a circular orbit of radius r about a nucleus. A uniform magnetic field \vec{B} is then established perpendicular to the plane of the orbit. Assuming also that the radius of the orbit does not change and that the change in the speed of the electron due to field \vec{B} is small, find an expression for the change in the orbital magnetic dipole moment of the electron due to the field.

sec. 32-10 Paramagnetism

•39 A sample of the paramagnetic salt to which the magnetization curve of Fig. 32-14 applies is to be tested to see whether it obeys Curie's law. The sample is placed in a uniform 0.50 T magnetic field that remains constant throughout the experiment. The magnetization M is then measured at temperatures ranging from 10 to 300 K . Will it be found that Curie's law is valid under these conditions?

•40 A sample of the paramagnetic salt to which the magnetization curve of Fig. 32-14 applies is held at room temperature (300 K). At what applied magnetic field will the degree of magnetic saturation of the sample be (a) 50% and (b) 90% ? (c) Are these fields attainable in the laboratory?

•41 SSM ILW A magnet in the form of a cylindrical rod has a length of 5.00 cm and a diameter of 1.00 cm . It has a uniform magnetization of $5.30 \times 10^3 \text{ A/m}$. What is its magnetic dipole moment?

•42 A 0.50 T magnetic field is applied to a paramagnetic gas whose atoms have an intrinsic magnetic dipole moment of $1.0 \times 10^{-23} \text{ J/T}$. At what temperature will the mean kinetic energy of translation of the atoms equal the energy required to reverse such a dipole end for end in this magnetic field?

•43 An electron with kinetic energy K_e travels in a circular path that is perpendicular to a uniform magnetic field, which is in the positive direction of a z axis. The electron's motion is subject only to the force due to the field. (a) Show that the magnetic dipole moment of the electron due to its orbital motion has magnitude $\mu = K_e/B$ and that it is in the direction opposite that of \vec{B} . What are the (b) magnitude and (c) direction of the magnetic dipole moment of a positive ion with kinetic energy K_i under the same circumstances? (d) An ionized gas consists of $5.3 \times 10^{21} \text{ electrons/m}^3$ and the same number density of ions. Take the average electron kinetic energy to be $6.2 \times 10^{-20} \text{ J}$ and the average ion kinetic energy to be $7.6 \times 10^{-21} \text{ J}$. Calculate the magnetization of the gas when it is in a magnetic field of 1.2 T .

•44 Figure 32-38 gives the magnetization curve for a paramagnetic material. The vertical axis scale is set by $a = 0.15$, and the horizontal axis scale is set by $b = 0.2 \text{ T/K}$. Let μ_{sam} be the measured net magnetic moment of a sample of the material and μ_{max} be the maximum possible net magnetic moment of that sample. According to Curie's law, what would be the ratio $\mu_{\text{sam}}/\mu_{\text{max}}$ were the sample placed in a uniform magnetic field of magnitude 0.800 T , at a temperature of 2.00 K ?

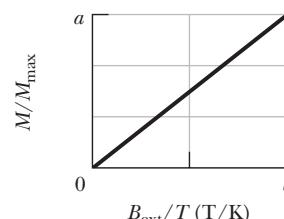


Fig. 32-38 Problem 44.

•45 SSM Consider a solid containing N atoms per unit volume, each atom having a magnetic dipole moment $\vec{\mu}$. Suppose the direction of $\vec{\mu}$ can be only parallel or antiparallel to an externally ap-

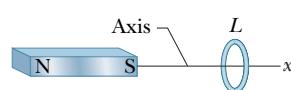


Fig. 32-37
Problems 37 and 71.

plied magnetic field \vec{B} (this will be the case if $\vec{\mu}$ is due to the spin of a single electron). According to statistical mechanics, the probability of an atom being in a state with energy U is proportional to $e^{-U/kT}$, where T is the temperature and k is Boltzmann's constant. Thus, because energy U is $-\vec{\mu} \cdot \vec{B}$, the fraction of atoms whose dipole moment is parallel to \vec{B} is proportional to $e^{\mu B/kT}$ and the fraction of atoms whose dipole moment is antiparallel to \vec{B} is proportional to $e^{-\mu B/kT}$. (a) Show that the magnitude of the magnetization of this solid is $M = N\mu \tanh(\mu B/kT)$. Here \tanh is the hyperbolic tangent function: $\tanh(x) = (e^x - e^{-x})/(e^x + e^{-x})$. (b) Show that the result given in (a) reduces to $M = N\mu^2 B/kT$ for $\mu B \ll kT$. (c) Show that the result of (a) reduces to $M = N\mu$ for $\mu B \gg kT$. (d) Show that both (b) and (c) agree qualitatively with Fig. 32-14.

sec. 32-11 Ferromagnetism

••46 GO You place a magnetic compass on a horizontal surface, allow the needle to settle, and then give the compass a gentle wiggle to cause the needle to oscillate about its equilibrium position. The oscillation frequency is 0.312 Hz. Earth's magnetic field at the location of the compass has a horizontal component of $18.0 \mu\text{T}$. The needle has a magnetic moment of 0.680 mJ/T . What is the needle's rotational inertia about its (vertical) axis of rotation?

••47 SSM ILW WWW The magnitude of the magnetic dipole moment of Earth is $8.0 \times 10^{22} \text{ J/T}$. (a) If the origin of this magnetism were a magnetized iron sphere at the center of Earth, what would be its radius? (b) What fraction of the volume of Earth would such a sphere occupy? Assume complete alignment of the dipoles. The density of Earth's inner core is 14 g/cm^3 . The magnetic dipole moment of an iron atom is $2.1 \times 10^{-23} \text{ J/T}$. (Note: Earth's inner core is in fact thought to be in both liquid and solid forms and partly iron, but a permanent magnet as the source of Earth's magnetism has been ruled out by several considerations. For one, the temperature is certainly above the Curie point.)

••48 The magnitude of the dipole moment associated with an atom of iron in an iron bar is $2.1 \times 10^{-23} \text{ J/T}$. Assume that all the atoms in the bar, which is 5.0 cm long and has a cross-sectional area of 1.0 cm^2 , have their dipole moments aligned. (a) What is the dipole moment of the bar? (b) What torque must be exerted to hold this magnet perpendicular to an external field of magnitude 1.5 T? (The density of iron is 7.9 g/cm^3 .)

••49 SSM The exchange coupling mentioned in Section 32-11 as being responsible for ferromagnetism is *not* the mutual magnetic interaction between two elementary magnetic dipoles. To show this, calculate (a) the magnitude of the magnetic field a distance of 10 nm away, along the dipole axis, from an atom with magnetic dipole moment $1.5 \times 10^{-23} \text{ J/T}$ (cobalt), and (b) the minimum energy required to turn a second identical dipole end for end in this field. (c) By comparing the latter with the mean translational kinetic energy of 0.040 eV, what can you conclude?

••50 A magnetic rod with length 6.00 cm, radius 3.00 mm, and (uniform) magnetization $2.70 \times 10^3 \text{ A/m}$ can turn about its center like a compass needle. It is placed in a uniform magnetic field \vec{B} of magnitude 35.0 mT, such that the directions of its dipole moment and \vec{B} make an angle of 68.0° . (a) What is the magnitude of the torque on the rod due to \vec{B} ? (b) What is the change in the orientation energy of the rod if the angle changes to 34.0° ?

••51 The saturation magnetization M_{\max} of the ferromagnetic metal nickel is $4.70 \times 10^5 \text{ A/m}$. Calculate the magnetic dipole moment of a single nickel atom. (The density of nickel is 8.90 g/cm^3 , and its molar mass is 58.71 g/mol.)

••52 Measurements in mines and boreholes indicate that Earth's interior temperature increases with depth at the average rate of $30 \text{ }^\circ\text{C}/\text{km}$. Assuming a surface temperature of 10°C , at what depth does iron cease to be ferromagnetic? (The Curie temperature of iron varies very little with pressure.)

••53 A Rowland ring is formed of ferromagnetic material. It is circular in cross section, with an inner radius of 5.0 cm and an outer radius of 6.0 cm, and is wound with 400 turns of wire. (a) What current must be set up in the windings to attain a toroidal field of magnitude $B_0 = 0.20 \text{ mT}$? (b) A secondary coil wound around the toroid has 50 turns and resistance 8.0Ω . If, for this value of B_0 , we have $B_M = 800B_0$, how much charge moves through the secondary coil when the current in the toroid windings is turned on?

Additional Problems

54 Using the approximations given in Problem 61, find (a) the altitude above Earth's surface where the magnitude of its magnetic field is 50.0% of the surface value at the same latitude; (b) the maximum magnitude of the magnetic field at the core–mantle boundary, 2900 km below Earth's surface; and the (c) magnitude and (d) inclination of Earth's magnetic field at the north geographic pole. (e) Suggest why the values you calculated for (c) and (d) differ from measured values.

55 Earth has a magnetic dipole moment of $8.0 \times 10^{22} \text{ J/T}$. (a) What current would have to be produced in a single turn of wire extending around Earth at its geomagnetic equator if we wished to set up such a dipole? Could such an arrangement be used to cancel out Earth's magnetism (b) at points in space well above Earth's surface or (c) on Earth's surface?

56 A charge q is distributed uniformly around a thin ring of radius r . The ring is rotating about an axis through its center and perpendicular to its plane, at an angular speed ω . (a) Show that the magnetic moment due to the rotating charge has magnitude $\mu = \frac{1}{2}q\omega r^2$. (b) What is the direction of this magnetic moment if the charge is positive?

57 A magnetic compass has its needle, of mass 0.050 kg and length 4.0 cm, aligned with the horizontal component of Earth's magnetic field at a place where that component has the value $B_h = 16 \mu\text{T}$. After the compass is given a momentary gentle shake, the needle oscillates with angular frequency $\omega = 45 \text{ rad/s}$. Assuming that the needle is a uniform thin rod mounted at its center, find the magnitude of its magnetic dipole moment.

58 The capacitor in Fig. 32-7 is being charged with a 2.50 A current. The wire radius is 1.50 mm, and the plate radius is 2.00 cm. Assume that the current i in the wire and the displacement current i_d in the capacitor gap are both uniformly distributed. What is the magnitude of the magnetic field due to i at the following radial distances from the wire's center: (a) 1.00 mm (inside the wire), (b) 3.00 mm (outside the wire), and (c) 2.20 cm (outside the wire)? What is the magnitude of the magnetic field due to i_d at the following radial distances from the central axis between the plates: (d) 1.00 mm (inside the gap), (e) 3.00 mm (inside the gap), and (f) 2.20 cm (outside the gap)? (g) Explain why the fields at the two smaller radii are so different for the wire and the gap but the fields at the largest radius are not.

59 A parallel-plate capacitor with circular plates of radius $R = 16 \text{ mm}$ and gap width $d = 5.0 \text{ mm}$ has a uniform electric field between the plates. Starting at time $t = 0$, the potential difference between the two plates is $V = (100 \text{ V})e^{-t/\tau}$, where the

time constant $\tau = 12 \text{ ms}$. At radial distance $r = 0.80R$ from the central axis, what is the magnetic field magnitude (a) as a function of time for $t \geq 0$ and (b) at time $t = 3\tau$?

- 60** A magnetic flux of 7.0 mWb is directed outward through the flat bottom face of the closed surface shown in Fig. 32-39. Along the flat top face (which has a radius of 4.2 cm) there is a 0.40 T magnetic field \vec{B} directed perpendicular to the face. What are the (a) magnitude and (b) direction (inward or outward) of the magnetic flux through the curved part of the surface?

- 61 SSM** The magnetic field of Earth can be approximated as the magnetic field of a dipole. The horizontal and vertical components of this field at any distance r from Earth's center are given by

$$B_h = \frac{\mu_0 \mu}{4\pi r^3} \cos \lambda_m, \quad B_v = \frac{\mu_0 \mu}{2\pi r^3} \sin \lambda_m,$$

where λ_m is the *magnetic latitude* (this type of latitude is measured from the geomagnetic equator toward the north or south geomagnetic pole). Assume that Earth's magnetic dipole moment has magnitude $\mu = 8.00 \times 10^{22} \text{ A} \cdot \text{m}^2$. (a) Show that the magnitude of Earth's field at latitude λ_m is given by

$$B = \frac{\mu_0 \mu}{4\pi r^3} \sqrt{1 + 3 \sin^2 \lambda_m}.$$

- (b) Show that the inclination ϕ_i of the magnetic field is related to the magnetic latitude λ_m by $\tan \phi_i = 2 \tan \lambda_m$.

- 62** Use the results displayed in Problem 61 to predict the (a) magnitude and (b) inclination of Earth's magnetic field at the geomagnetic equator, the (c) magnitude and (d) inclination at geomagnetic latitude 60.0° , and the (e) magnitude and (f) inclination at the north geomagnetic pole.

- 63** A parallel-plate capacitor with circular plates of radius 55.0 mm is being charged. At what radius (a) inside and (b) outside the capacitor gap is the magnitude of the induced magnetic field equal to 50.0% of its maximum value?

- 64** A sample of the paramagnetic salt to which the magnetization curve of Fig. 32-14 applies is immersed in a uniform magnetic field of 2.0 T . At what temperature will the degree of magnetic saturation of the sample be (a) 50% and (b) 90%?

- 65** A parallel-plate capacitor with circular plates of radius R is being discharged. The displacement current through a central circular area, parallel to the plates and with radius $R/2$, is 2.0 A . What is the discharging current?

- 66** Figure 32-40 gives the variation of an electric field that is perpendicular to a circular area of 2.0 m^2 . During the time period shown, what is the greatest displacement current through the area?

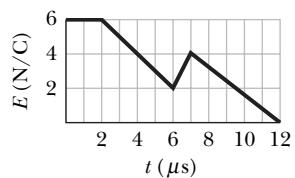


Fig. 32-40 Problem 66.

- 67** In Fig. 32-41, a parallel-plate capacitor is being discharged by a current $i = 5.0 \text{ A}$. The plates are square with edge length $L = 8.0$

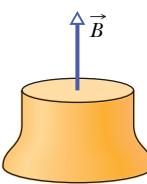


Fig. 32-39
Problem 60.

mm. (a) What is the rate at which the electric field between the plates is changing? (b) What is the value of $\oint \vec{B} \cdot d\vec{s}$ around the dashed path, where $H = 2.0 \text{ mm}$ and $W = 3.0 \text{ mm}$?

- 68** What is the measured component of the orbital magnetic dipole moment of an electron with the values (a) $m_\ell = 3$ and (b) $m_\ell = -4$?

- 69** In Fig. 32-42, a bar magnet lies near a paper cylinder. (a) Sketch the magnetic field lines that pass through the surface of the cylinder. (b) What is the sign of $\vec{B} \cdot d\vec{A}$ for every area $d\vec{A}$ on the surface? (c) Does this contradict Gauss' law for magnetism? Explain.

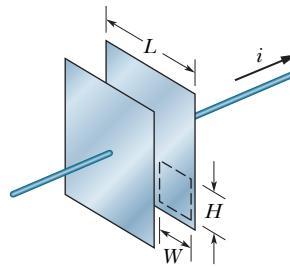


Fig. 32-41 Problem 67.



Fig. 32-42 Problem 69.

- 70** In the lowest energy state of the hydrogen atom, the most probable distance of the single electron from the central proton (the nucleus) is $5.2 \times 10^{-11} \text{ m}$. (a) Compute the magnitude of the proton's electric field at that distance. The component $\mu_{s,z}$ of the proton's spin magnetic dipole moment measured on a z axis is $1.4 \times 10^{-26} \text{ J/T}$. (b) Compute the magnitude of the proton's magnetic field at the distance $5.2 \times 10^{-11} \text{ m}$ on the z axis. (Hint: Use Eq. 29-27.) (c) What is the ratio of the spin magnetic dipole moment of the electron to that of the proton?

- 71** Figure 32-37 shows a loop model (loop L) for a paramagnetic material. (a) Sketch the field lines through and about the material due to the magnet. What is the direction of (b) the loop's net magnetic dipole moment $\vec{\mu}$, (c) the conventional current i in the loop (clockwise or counterclockwise in the figure), and (d) the magnetic force acting on the loop?

- 72** Two plates (as in Fig. 32-7) are being discharged by a constant current. Each plate has a radius of 4.00 cm . During the discharging, at a point between the plates at radial distance 2.00 cm from the central axis, the magnetic field has a magnitude of 12.5 nT . (a) What is the magnitude of the magnetic field at radial distance 6.00 cm ? (b) What is the current in the wires attached to the plates?

- 73 SSM** If an electron in an atom has orbital angular momentum with m_ℓ values limited by ± 3 , how many values of (a) $L_{\text{orb},z}$ and (b) $\mu_{\text{orb},z}$ can the electron have? In terms of h, m , and e , what is the greatest allowed magnitude for (c) $L_{\text{orb},z}$ and (d) $\mu_{\text{orb},z}$? (e) What is the greatest allowed magnitude for the z component of the electron's *net* angular momentum (orbital plus spin)? (f) How many values (signs included) are allowed for the z component of its *net* angular momentum?

- 74** A parallel-plate capacitor with circular plates is being charged. Consider a circular loop centered on the central axis and located between the plates. If the loop radius of 3.00 cm is greater than the plate radius, what is the displacement current between the plates when the magnetic field along the loop has magnitude $2.00 \mu\text{T}$?

- 75** Suppose that ± 4 are the limits to the values of m_ℓ for an electron in an atom. (a) How many different values of the electron's $\mu_{\text{orb},z}$ are possible? (b) What is the greatest magnitude of those possible values? Next, if the atom is in a magnetic field of magnitude 0.250 T , in the positive direction of the z axis, what are (c) the maximum energy and (d) the minimum energy associated with those possible values of $\mu_{\text{orb},z}$?

ELECTROMAGNETIC WAVES

33

33-1 WHAT IS PHYSICS?

The information age in which we live is based almost entirely on the physics of electromagnetic waves. Like it or not, we are now globally connected by television, telephones, and the Web. And like it or not, we are constantly immersed in those signals because of television, radio, and telephone transmitters.

Much of this global interconnection of information processors was not imagined by even the most visionary engineers of 20 years ago. The challenge for today's engineers is trying to envision what the global interconnection will be like 20 years from now. The starting point in meeting that challenge is understanding the basic physics of electromagnetic waves, which come in so many different types that they are poetically said to form *Maxwell's rainbow*.

33-2 Maxwell's Rainbow

The crowning achievement of James Clerk Maxwell (see Chapter 32) was to show that a beam of light is a traveling wave of electric and magnetic fields—an **electromagnetic wave**—and thus that optics, the study of visible light, is a branch of electromagnetism. In this chapter we move from one to the other: we conclude our discussion of strictly electrical and magnetic phenomena, and we build a foundation for optics.

In Maxwell's time (the mid 1800s), the visible, infrared, and ultraviolet forms of light were the only electromagnetic waves known. Spurred on by Maxwell's work, however, Heinrich Hertz discovered what we now call radio waves and verified that they move through the laboratory at the same speed as visible light.

As Fig. 33-1 shows, we now know a wide *spectrum* (or range) of electromagnetic waves: Maxwell's rainbow. Consider the extent to which we are immersed in electromagnetic waves throughout this spectrum. The Sun, whose radiations define the environment in which we as a species have evolved and adapted, is the dominant source. We are also crisscrossed by radio and television signals. Microwaves from radar systems and from telephone relay systems may reach us. There are electromagnetic waves from lightbulbs, from the heated engine blocks of automobiles, from x-ray machines, from lightning flashes, and from buried radioactive materials. Beyond this, radiation reaches us from stars and other objects in our galaxy and from other galaxies. Electromagnetic waves also travel in the other direction. Television signals, transmitted from Earth since about 1950, have now taken news about us (along with episodes of *I Love Lucy*, albeit *very* faintly) to whatever technically sophisticated inhabitants there may be on whatever planets may encircle the nearest 400 or so stars.

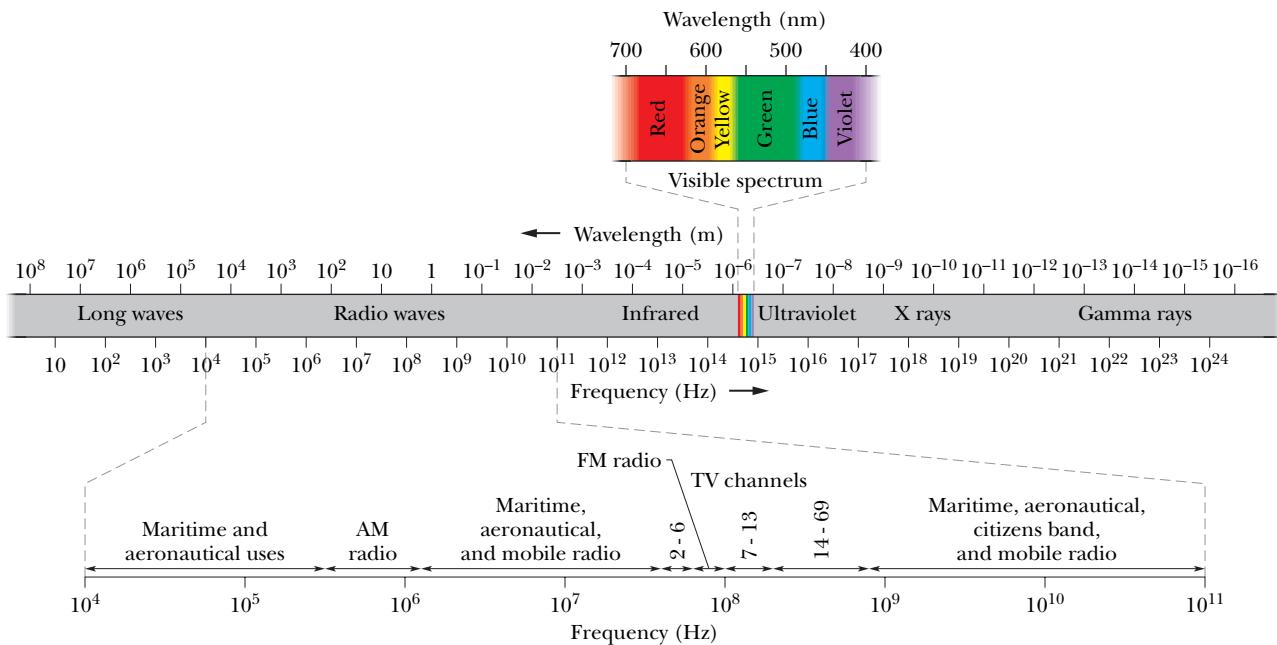


Fig. 33-1 The electromagnetic spectrum.

In the wavelength scale in Fig. 33-1 (and similarly the corresponding frequency scale), each scale marker represents a change in wavelength (and correspondingly in frequency) by a factor of 10. The scale is open-ended; the wavelengths of electromagnetic waves have no inherent upper or lower bound.

Certain regions of the electromagnetic spectrum in Fig. 33-1 are identified by familiar labels, such as *x rays* and *radio waves*. These labels denote roughly defined wavelength ranges within which certain kinds of sources and detectors of electromagnetic waves are in common use. Other regions of Fig. 33-1, such as those labeled *TV channels* and *AM radio*, represent specific wavelength bands assigned by law for certain commercial or other purposes. There are no gaps in the electromagnetic spectrum—and all electromagnetic waves, no matter where they lie in the spectrum, travel through *free space* (vacuum) with the same speed c .

The visible region of the spectrum is of course of particular interest to us. Figure 33-2 shows the relative sensitivity of the human eye to light of various wavelengths. The center of the visible region is about 555 nm, which produces the sensation that we call yellow-green.

The limits of this visible spectrum are not well defined because the eye sensitivity curve approaches the zero-sensitivity line asymptotically at both long and short wavelengths. If we take the limits, arbitrarily, as the wavelengths at which eye sensitivity has dropped to 1% of its maximum value, these limits are about 430 and 690 nm; however, the eye can detect electromagnetic waves somewhat beyond these limits if they are intense enough.

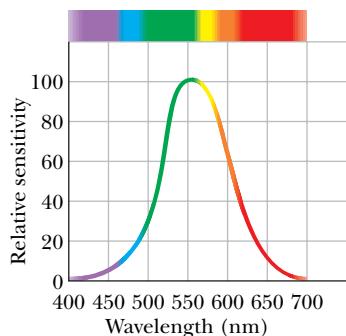


Fig. 33-2 The relative sensitivity of the average human eye to electromagnetic waves at different wavelengths. This portion of the electromagnetic spectrum to which the eye is sensitive is called *visible light*.

33-3 The Traveling Electromagnetic Wave, Qualitatively

Some electromagnetic waves, including *x rays*, *gamma rays*, and *visible light*, are *radiated* (emitted) from sources that are of atomic or nuclear size, where quantum physics rules. Here we discuss how other electromagnetic waves are generated. To simplify matters, we restrict ourselves to that region of the spectrum (wavelength $\lambda \approx 1$ m) in which the source of the *radiation* (the emitted waves) is both macroscopic and of manageable dimensions.

33-3 THE TRAVELING ELECTROMAGNETIC WAVE, QUALITATIVELY

891

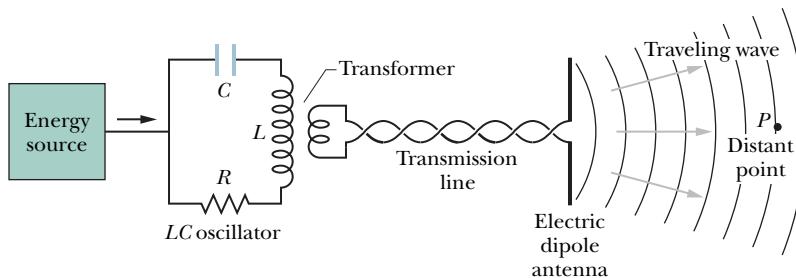


Fig. 33-3 An arrangement for generating a traveling electromagnetic wave in the shortwave radio region of the spectrum: an *LC oscillator* produces a sinusoidal current in the antenna, which generates the wave. *P* is a distant point at which a detector can monitor the wave traveling past it.

Figure 33-3 shows, in broad outline, the generation of such waves. At its heart is an *LC oscillator*, which establishes an angular frequency ω ($= 1/\sqrt{LC}$). Charges and currents in this circuit vary sinusoidally at this frequency, as depicted in Fig. 31-1. An external source—possibly an ac generator—must be included to supply energy to compensate both for thermal losses in the circuit and for energy carried away by the radiated electromagnetic wave.

The *LC oscillator* of Fig. 33-3 is coupled by a transformer and a transmission line to an *antenna*, which consists essentially of two thin, solid, conducting rods. Through this coupling, the sinusoidally varying current in the oscillator causes charge to oscillate sinusoidally along the rods of the antenna at the angular frequency ω of the *LC oscillator*. The current in the rods associated with this movement of charge also varies sinusoidally, in magnitude and direction, at angular frequency ω . The antenna has the effect of an electric dipole whose electric dipole moment varies sinusoidally in magnitude and direction along the antenna.

Because the dipole moment varies in magnitude and direction, the electric field produced by the dipole varies in magnitude and direction. Also, because the current varies, the magnetic field produced by the current varies in magnitude and direction. However, the changes in the electric and magnetic fields do not happen everywhere instantaneously; rather, the changes travel outward from the antenna at the speed of light c . Together the changing fields form an electromagnetic wave that travels away from the antenna at speed c . The angular frequency of this wave is ω , the same as that of the *LC oscillator*.

Figure 33-4 shows how the electric field \vec{E} and the magnetic field \vec{B} change with time as one wavelength of the wave sweeps past the distant point *P* of Fig. 33-3; in each part of Fig. 33-4, the wave is traveling directly out of the page. (We choose a distant point so that the curvature of the waves suggested in Fig. 33-3 is small enough to neglect. At such points, the wave is said to be a *plane wave*, and discussion of the wave is much simplified.) Note several key features in Fig. 33-4; they are present regardless of how the wave is created:

1. The electric and magnetic fields \vec{E} and \vec{B} are always perpendicular to the direction in which the wave is traveling. Thus, the wave is a *transverse wave*, as discussed in Chapter 16.
2. The electric field is always perpendicular to the magnetic field.
3. The cross product $\vec{E} \times \vec{B}$ always gives the direction in which the wave travels.
4. The fields always vary sinusoidally, just like the transverse waves discussed in Chapter 16. Moreover, the fields vary with the same frequency and *in phase* (in step) with each other.

In keeping with these features, we can assume that the electromagnetic wave is traveling toward *P* in the positive direction of an *x* axis, that the electric field in

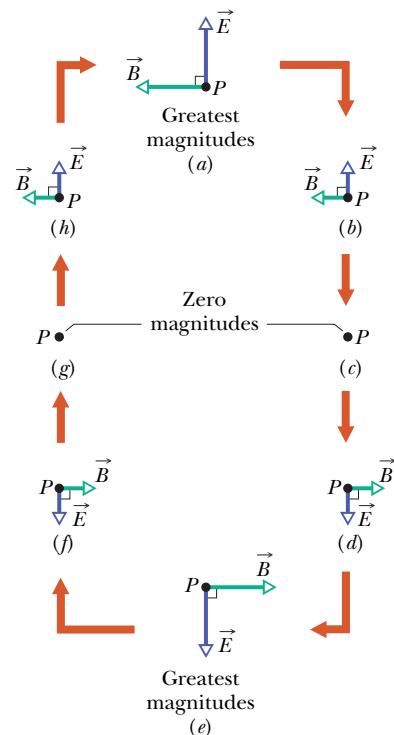


Fig. 33-4 (a)–(h) The variation in the electric field \vec{E} and the magnetic field \vec{B} at the distant point *P* of Fig. 33-3 as one wavelength of the electromagnetic wave travels past it. In this perspective, the wave is traveling directly out of the page. The two fields vary sinusoidally in magnitude and direction. Note that they are always perpendicular to each other and to the wave's direction of travel.

Fig. 33-4 is oscillating parallel to the y axis, and that the magnetic field is then oscillating parallel to the z axis (using a right-handed coordinate system, of course). Then we can write the electric and magnetic fields as sinusoidal functions of position x (along the path of the wave) and time t :

$$E = E_m \sin(kx - \omega t), \quad (33-1)$$

$$B = B_m \sin(kx - \omega t), \quad (33-2)$$

in which E_m and B_m are the amplitudes of the fields and, as in Chapter 16, ω and k are the angular frequency and angular wave number of the wave, respectively. From these equations, we note that not only do the two fields form the electromagnetic wave but each also forms its own wave. Equation 33-1 gives the *electric wave component* of the electromagnetic wave, and Eq. 33-2 gives the *magnetic wave component*. As we shall discuss below, these two wave components cannot exist independently.

From Eq. 16-13, we know that the speed of the wave is ω/k . However, because this is an electromagnetic wave, its speed (in vacuum) is given by the symbol c rather than v . In the next section you will see that c has the value

$$c = \frac{1}{\sqrt{\mu_0 \epsilon_0}} \quad (\text{wave speed}), \quad (33-3)$$

which is about 3.0×10^8 m/s. In other words,



All electromagnetic waves, including visible light, have the same speed c in vacuum.

You will also see that the wave speed c and the amplitudes of the electric and magnetic fields are related by

$$\frac{E_m}{B_m} = c \quad (\text{amplitude ratio}). \quad (33-4)$$

If we divide Eq. 33-1 by Eq. 33-2 and then substitute with Eq. 33-4, we find that the magnitudes of the fields at every instant and at any point are related by

$$\frac{E}{B} = c \quad (\text{magnitude ratio}). \quad (33-5)$$

We can represent the electromagnetic wave as in Fig. 33-5a, with a *ray* (a directed line showing the wave's direction of travel) or with *wavefronts* (imaginary surfaces over which the wave has the same magnitude of electric field), or both. The two wavefronts shown in Fig. 33-5a are separated by one wavelength $\lambda (= 2\pi/k)$ of the wave. (Waves traveling in approximately the same direction form a *beam*, such as a laser beam, which can also be represented with a ray.)

We can also represent the wave as in Fig. 33-5b, which shows the electric and magnetic field vectors in a “snapshot” of the wave at a certain instant. The curves through the tips of the vectors represent the sinusoidal oscillations given by Eqs. 33-1 and 33-2; the wave components \vec{E} and \vec{B} are in phase, perpendicular to each other, and perpendicular to the wave's direction of travel.

Interpretation of Fig. 33-5b requires some care. The similar drawings for a transverse wave on a taut string that we discussed in Chapter 16 represented the up and down displacement of sections of the string as the wave passed (*something actually moved*). Figure 33-5b is more abstract. At the instant shown, the electric and magnetic fields each have a certain magnitude and direction (but always perpendicular to the x axis) at each point along the x axis. We choose to represent these vector

33-3 THE TRAVELING ELECTROMAGNETIC WAVE, QUALITATIVELY

893

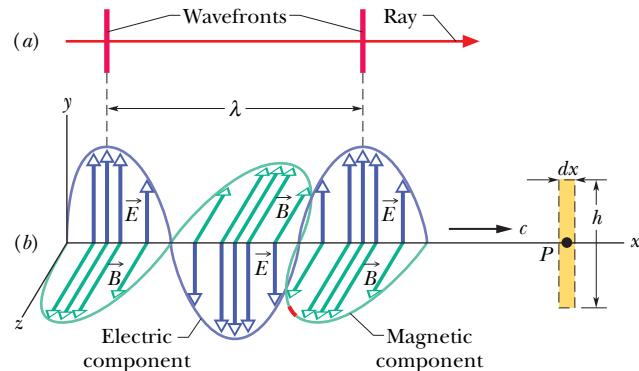


Fig. 33-5 (a) An electromagnetic wave represented with a ray and two wavefronts; the wavefronts are separated by one wavelength λ . (b) The same wave represented in a “snapshot” of its electric field \vec{E} and magnetic field \vec{B} at points on the x axis, along which the wave travels at speed c . As it travels past point P , the fields vary as shown in Fig. 33-4. The electric component of the wave consists of only the electric fields; the magnetic component consists of only the magnetic fields. The dashed rectangle at P is used in Fig. 33-6.

quantities with a pair of arrows for each point, and so we must draw arrows of different lengths for different points, all directed away from the x axis, like thorns on a rose stem. However, the arrows represent field values only at points that are on the x axis. Neither the arrows nor the sinusoidal curves represent a sideways motion of anything, nor do the arrows connect points on the x axis with points off the axis.

Drawings like Fig. 33-5 help us visualize what is actually a very complicated situation. First consider the magnetic field. Because it varies sinusoidally, it induces (via Faraday’s law of induction) a perpendicular electric field that also varies sinusoidally. However, because that electric field is varying sinusoidally, it induces (via Maxwell’s law of induction) a perpendicular magnetic field that also varies sinusoidally. And so on. The two fields continuously create each other via induction, and the resulting sinusoidal variations in the fields travel as a wave—the electromagnetic wave. Without this amazing result, we could not see; indeed, because we need electromagnetic waves from the Sun to maintain Earth’s temperature, without this result we could not even exist.

A Most Curious Wave

The waves we discussed in Chapters 16 and 17 require a *medium* (some material) through which or along which to travel. We had waves traveling along a string, through Earth, and through the air. However, an electromagnetic wave (let’s use the term *light wave* or *light*) is curiously different in that it requires no medium for its travel. It can, indeed, travel through a medium such as air or glass, but it can also travel through the vacuum of space between a star and us.

Once the special theory of relativity became accepted, long after Einstein published it in 1905, the speed of light waves was realized to be special. One reason is that light has the same speed regardless of the frame of reference from which it is measured. If you send a beam of light along an axis and ask several observers to measure its speed while they move at different speeds along that axis, either in the direction of the light or opposite it, they will all measure the *same speed* for the light. This result is an amazing one and quite different from what would have been found if those observers had measured the speed of any other type of wave; for other waves, the speed of the observers relative to the wave would have affected their measurements.

The meter has now been defined so that the speed of light (any electromagnetic wave) in vacuum has the exact value

$$c = 299\,792\,458 \text{ m/s},$$

which can be used as a standard. In fact, if you now measure the travel time of a pulse of light from one point to another, you are not really measuring the speed of the light but rather the distance between those two points.

The oscillating magnetic field induces an oscillating and perpendicular electric field.

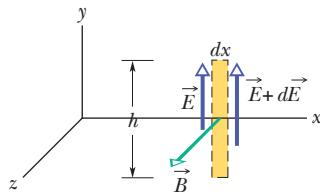


Fig. 33-6 As the electromagnetic wave travels rightward past point P in Fig. 33-5b, the sinusoidal variation of the magnetic field \vec{B} through a rectangle centered at P induces electric fields along the rectangle. At the instant shown, \vec{B} is decreasing in magnitude and the induced electric field is therefore greater in magnitude on the right side of the rectangle than on the left.

33-4 The Traveling Electromagnetic Wave, Quantitatively

We shall now derive Eqs. 33-3 and 33-4 and, even more important, explore the dual induction of electric and magnetic fields that gives us light.

Equation 33-4 and the Induced Electric Field

The dashed rectangle of dimensions dx and h in Fig. 33-6 is fixed at point P on the x axis and in the xy plane (it is shown on the right in Fig. 33-5b). As the electromagnetic wave moves rightward past the rectangle, the magnetic flux Φ_B through the rectangle changes and—according to Faraday's law of induction—induced electric fields appear throughout the region of the rectangle. We take \vec{E} and $\vec{E} + d\vec{E}$ to be the induced fields along the two long sides of the rectangle. These induced electric fields are, in fact, the electrical component of the electromagnetic wave.

Note the small red portion of the magnetic field component curve far from the y axis in Fig. 33-5b. Let's consider the induced electric fields at the instant when this red portion of the magnetic component is passing through the rectangle. Just then, the magnetic field through the rectangle points in the positive z direction and is decreasing in magnitude (the magnitude was greater just before the red section arrived). Because the magnetic field is decreasing, the magnetic flux Φ_B through the rectangle is also decreasing. According to Faraday's law, this change in flux is opposed by induced electric fields, which produce a magnetic field \vec{B} in the positive z direction.

According to Lenz's law, this in turn means that if we imagine the boundary of the rectangle to be a conducting loop, a counterclockwise induced current would have to appear in it. There is, of course, no conducting loop; but this analysis shows that the induced electric field vectors \vec{E} and $\vec{E} + d\vec{E}$ are indeed oriented as shown in Fig. 33-6, with the magnitude of $\vec{E} + d\vec{E}$ greater than that of \vec{E} . Otherwise, the net induced electric field would not act counterclockwise around the rectangle.

Let us now apply Faraday's law of induction,

$$\oint \vec{E} \cdot d\vec{s} = -\frac{d\Phi_B}{dt}, \quad (33-6)$$

counterclockwise around the rectangle of Fig. 33-6. There is no contribution to the integral from the top or bottom of the rectangle because \vec{E} and $d\vec{s}$ are perpendicular to each other there. The integral then has the value

$$\oint \vec{E} \cdot d\vec{s} = (E + dE)h - Eh = h dE. \quad (33-7)$$

The flux Φ_B through this rectangle is

$$\Phi_B = (B)(h dx), \quad (33-8)$$

where B is the average magnitude of \vec{B} within the rectangle and $h dx$ is the area of the rectangle. Differentiating Eq. 33-8 with respect to t gives

$$\frac{d\Phi_B}{dt} = h dx \frac{dB}{dt}. \quad (33-9)$$

If we substitute Eqs. 33-7 and 33-9 into Eq. 33-6, we find

$$h dE = -h dx \frac{dB}{dt}$$

or

$$\frac{dE}{dx} = -\frac{dB}{dt}. \quad (33-10)$$

33-4 THE TRAVELING ELECTROMAGNETIC WAVE, QUANTITATIVELY

895

Actually, both B and E are functions of *two* variables, x and t , as Eqs. 33-1 and 33-2 show. However, in evaluating dE/dx , we must assume that t is constant because Fig. 33-6 is an “instantaneous snapshot.” Also, in evaluating dB/dt we must assume that x is constant because we are dealing with the time rate of change of B at a particular place, the point P in Fig. 33-5b. The derivatives under these circumstances are *partial derivatives*, and Eq. 33-10 must be written

$$\frac{\partial E}{\partial x} = -\frac{\partial B}{\partial t}. \quad (33-11)$$

The minus sign in this equation is appropriate and necessary because, although E is increasing with x at the site of the rectangle in Fig. 33-6, B is decreasing with t .

From Eq. 33-1 we have

$$\frac{\partial E}{\partial x} = kE_m \cos(kx - \omega t)$$

and from Eq. 33-2

$$\frac{\partial B}{\partial t} = -\omega B_m \cos(kx - \omega t).$$

Then Eq. 33-11 reduces to

$$kE_m \cos(kx - \omega t) = \omega B_m \cos(kx - \omega t). \quad (33-12)$$

The ratio ω/k for a traveling wave is its speed, which we are calling c . Equation 33-12 then becomes

$$\frac{E_m}{B_m} = c \quad (\text{amplitude ratio}), \quad (33-13)$$

which is just Eq. 33-4.

Equation 33-3 and the Induced Magnetic Field

Figure 33-7 shows another dashed rectangle at point P of Fig. 33-5b; this one is in the xz plane. As the electromagnetic wave moves rightward past this new rectangle, the electric flux Φ_E through the rectangle changes and—according to Maxwell’s law of induction—induced magnetic fields appear throughout the region of the rectangle. These induced magnetic fields are, in fact, the magnetic component of the electromagnetic wave.

We see from Fig. 33-5b that at the instant chosen for the magnetic field represented in Fig. 33-6, marked in red on the magnetic component curve, the electric field through the rectangle of Fig. 33-7 is directed as shown. Recall that at the chosen instant, the magnetic field in Fig. 33-6 is decreasing. Because the two fields are in phase, the electric field in Fig. 33-7 must also be decreasing, and so must the electric flux Φ_E through the rectangle. By applying the same reasoning we applied to Fig. 33-6, we see that the changing flux Φ_E will induce a magnetic field with vectors \vec{B} and $\vec{B} + d\vec{B}$ oriented as shown in Fig. 33-7, where field $\vec{B} + d\vec{B}$ is greater than field \vec{B} .

The oscillating electric field induces an oscillating and perpendicular magnetic field.

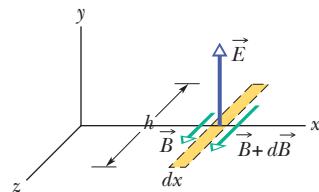


Fig. 33-7 The sinusoidal variation of the electric field through this rectangle, located (but not shown) at point P in Fig. 33-5b, induces magnetic fields along the rectangle. The instant shown is that of Fig. 33-6: \vec{E} is decreasing in magnitude, and the magnitude of the induced magnetic field is greater on the right side of the rectangle than on the left.

Let us apply Maxwell's law of induction,

$$\oint \vec{B} \cdot d\vec{s} = \mu_0 \epsilon_0 \frac{d\Phi_E}{dt}, \quad (33-14)$$

by proceeding counterclockwise around the dashed rectangle of Fig. 33-7. Only the long sides of the rectangle contribute to the integral because the dot product along the short sides is zero. Thus, we can write

$$\oint \vec{B} \cdot d\vec{s} = -(B + dB)h + Bh = -h dB. \quad (33-15)$$

The flux Φ_E through the rectangle is

$$\Phi_E = (E)(h dx), \quad (33-16)$$

where E is the average magnitude of \vec{E} within the rectangle. Differentiating Eq. 33-16 with respect to t gives

$$\frac{d\Phi_E}{dt} = h dx \frac{dE}{dt}.$$

If we substitute this and Eq. 33-15 into Eq. 33-14, we find

$$-h dB = \mu_0 \epsilon_0 \left(h dx \frac{dE}{dt} \right)$$

or, changing to partial-derivative notation as we did for Eq. 33-11,

$$-\frac{\partial B}{\partial x} = \mu_0 \epsilon_0 \frac{\partial E}{\partial t}. \quad (33-17)$$

Again, the minus sign in this equation is necessary because, although B is increasing with x at point P in the rectangle in Fig. 33-7, E is decreasing with t .

Evaluating Eq. 33-17 by using Eqs. 33-1 and 33-2 leads to

$$-kB_m \cos(kx - \omega t) = -\mu_0 \epsilon_0 \omega E_m \cos(kx - \omega t),$$

which we can write as

$$\frac{E_m}{B_m} = \frac{1}{\mu_0 \epsilon_0 (\omega/k)} = \frac{1}{\mu_0 \epsilon_0 c}.$$

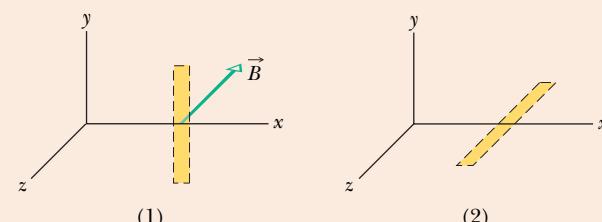
Combining this with Eq. 33-13 leads at once to

$$c = \frac{1}{\sqrt{\mu_0 \epsilon_0}} \quad (\text{wave speed}), \quad (33-18)$$

which is exactly Eq. 33-3.

CHECKPOINT 1

The magnetic field \vec{B} through the rectangle of Fig. 33-6 is shown at a different instant in part 1 of the figure here; \vec{B} is directed in the xz plane, parallel to the z axis, and its magnitude is increasing. (a) Complete part 1 by drawing the induced electric fields, indicating both directions and relative magnitudes (as in Fig. 33-6). (b) For the same instant, complete part 2 of the figure by drawing the electric field of the electromagnetic wave. Also draw the induced magnetic fields, indicating both directions and relative magnitudes (as in Fig. 33-7).



33-5 Energy Transport and the Poynting Vector

All sunbathers know that an electromagnetic wave can transport energy and deliver it to a body on which the wave falls. The rate of energy transport per unit area in such a wave is described by a vector \vec{S} , called the **Poynting vector** after physicist John Henry Poynting (1852–1914), who first discussed its properties. This vector is defined as

$$\vec{S} = \frac{1}{\mu_0} \vec{E} \times \vec{B} \quad (\text{Poynting vector}). \quad (33-19)$$

Its magnitude S is related to the rate at which energy is transported by a wave across a unit area at any instant (inst):

$$S = \left(\frac{\text{energy/time}}{\text{area}} \right)_{\text{inst}} = \left(\frac{\text{power}}{\text{area}} \right)_{\text{inst}}. \quad (33-20)$$

From this we can see that the SI unit for \vec{S} is the watt per square meter (W/m^2).



The direction of the Poynting vector \vec{S} of an electromagnetic wave at any point gives the wave's direction of travel and the direction of energy transport at that point.

Because \vec{E} and \vec{B} are perpendicular to each other in an electromagnetic wave, the magnitude of $\vec{E} \times \vec{B}$ is EB . Then the magnitude of \vec{S} is

$$S = \frac{1}{\mu_0} EB, \quad (33-21)$$

in which S , E , and B are instantaneous values. The magnitudes E and B are so closely coupled to each other that we need to deal with only one of them; we choose E , largely because most instruments for detecting electromagnetic waves deal with the electric component of the wave rather than the magnetic component. Using $B = E/c$ from Eq. 33-5, we can rewrite Eq. 33-21 in terms of just the electric component as

$$S = \frac{1}{c\mu_0} E^2 \quad (\text{instantaneous energy flow rate}). \quad (33-22)$$

By substituting $E = E_m \sin(kx - \omega t)$ into Eq. 33-22, we could obtain an equation for the energy transport rate as a function of time. More useful in practice, however, is the average energy transported over time; for that, we need to find the time-averaged value of S , written S_{avg} and also called the **intensity** I of the wave. Thus from Eq. 33-20, the intensity I is

$$I = S_{\text{avg}} = \left(\frac{\text{energy/time}}{\text{area}} \right)_{\text{avg}} = \left(\frac{\text{power}}{\text{area}} \right)_{\text{avg}}. \quad (33-23)$$

From Eq. 33-22, we find

$$I = S_{\text{avg}} = \frac{1}{c\mu_0} [E^2]_{\text{avg}} = \frac{1}{c\mu_0} [E_m^2 \sin^2(kx - \omega t)]_{\text{avg}}. \quad (33-24)$$

Over a full cycle, the average value of $\sin^2 \theta$, for any angular variable θ , is $\frac{1}{2}$ (see Fig. 31-17). In addition, we define a new quantity E_{rms} , the *root-mean-square* value of the electric field, as

$$E_{\text{rms}} = \frac{E_m}{\sqrt{2}}. \quad (33-25)$$

The energy emitted by light source S must pass through the sphere of radius r .

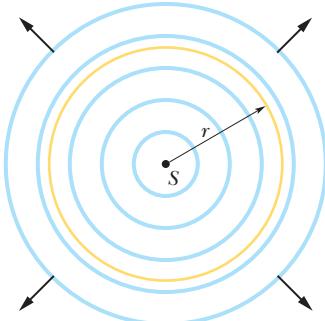


Fig. 33-8 A point source S emits electromagnetic waves uniformly in all directions. The spherical wavefronts pass through an imaginary sphere of radius r that is centered on S .

We can then rewrite Eq. 33-24 as

$$I = \frac{1}{c\mu_0} E_{\text{rms}}^2. \quad (33-26)$$

Because $E = cB$ and c is such a very large number, you might conclude that the energy associated with the electric field is much greater than that associated with the magnetic field. That conclusion is incorrect; the two energies are exactly equal. To show this, we start with Eq. 25-25, which gives the energy density u ($= \frac{1}{2}\epsilon_0 E^2$) within an electric field, and substitute cB for E ; then we can write

$$u_E = \frac{1}{2}\epsilon_0 E^2 = \frac{1}{2}\epsilon_0 (cB)^2.$$

If we now substitute for c with Eq. 33-3, we get

$$u_E = \frac{1}{2}\epsilon_0 \frac{1}{\mu_0\epsilon_0} B^2 = \frac{B^2}{2\mu_0}.$$

However, Eq. 30-55 tells us that $B^2/2\mu_0$ is the energy density u_B of a magnetic field \vec{B} ; so we see that $u_E = u_B$ everywhere along an electromagnetic wave.

Variation of Intensity with Distance

How intensity varies with distance from a real source of electromagnetic radiation is often complex—especially when the source (like a searchlight at a movie premier) beams the radiation in a particular direction. However, in some situations we can assume that the source is a *point source* that emits the light *isotropically*—that is, with equal intensity in all directions. The spherical wavefronts spreading from such an isotropic point source S at a particular instant are shown in cross section in Fig. 33-8.

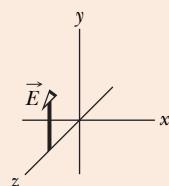
Let us assume that the energy of the waves is conserved as they spread from this source. Let us also center an imaginary sphere of radius r on the source, as shown in Fig. 33-8. All the energy emitted by the source must pass through the sphere. Thus, the rate at which energy passes through the sphere via the radiation must equal the rate at which energy is emitted by the source—that is, the source power P_s . The intensity I (power per unit area) measured at the sphere must then be, from Eq. 33-23,

$$I = \frac{\text{power}}{\text{area}} = \frac{P_s}{4\pi r^2}, \quad (33-27)$$

where $4\pi r^2$ is the area of the sphere. Equation 33-27 tells us that the intensity of the electromagnetic radiation from an isotropic point source decreases with the square of the distance r from the source.

CHECKPOINT 2

The figure here gives the electric field of an electromagnetic wave at a certain point and a certain instant. The wave is transporting energy in the negative z direction. What is the direction of the magnetic field of the wave at that point and instant?



Sample Problem**Light wave: rms values of the electric and magnetic fields**

When you look at the North Star (Polaris), you intercept light from a star at a distance of 431 ly and emitting energy at a rate of 2.2×10^3 times that of our Sun ($P_{\text{sun}} = 3.90 \times 10^{26}$ W). Neglecting any atmospheric absorption, find the rms values of the electric and magnetic fields when the starlight reaches you.

KEY IDEAS

1. The rms value E_{rms} of the electric field in light is related to the intensity I of the light via Eq. 33-26 ($I = E_{\text{rms}}^2/c\mu_0$).
2. Because the source is so far away and emits light with equal intensity in all directions, the intensity I at any distance r from the source is related to the source's power P_s via Eq. 33-27 ($I = P_s/4\pi r^2$).
3. The magnitudes of the electric field and magnetic field of an electromagnetic wave at any instant and at any point in the wave are related by the speed of light c according to Eq. 33-5 ($E/B = c$). Thus, the rms values of those fields are also related by Eq. 33-5.

Electric field: Putting the first two ideas together gives us

$$I = \frac{P_s}{4\pi r^2} = \frac{E_{\text{rms}}^2}{c\mu_0}$$



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and

$$E_{\text{rms}} = \sqrt{\frac{P_s c \mu_0}{4\pi r^2}}.$$

Substituting $P_s = (2.2 \times 10^3)(3.90 \times 10^{26}$ W), $r = 431$ ly = 4.08×10^{18} m, and values for the constants, we find

$$E_{\text{rms}} = 1.24 \times 10^{-3} \text{ V/m} \approx 1.2 \text{ mV/m.} \quad (\text{Answer})$$

Magnetic field: From Eq. 33-5, we write

$$\begin{aligned} B_{\text{rms}} &= \frac{E_{\text{rms}}}{c} = \frac{1.24 \times 10^{-3} \text{ V/m}}{3.00 \times 10^8 \text{ m/s}} \\ &= 4.1 \times 10^{-12} \text{ T} = 4.1 \text{ pT.} \end{aligned}$$

Cannot compare the fields: Note that E_{rms} (= 1.2 mV/m) is small as judged by ordinary laboratory standards, but B_{rms} (= 4.1 pT) is quite small. This difference helps to explain why most instruments used for the detection and measurement of electromagnetic waves are designed to respond to the electric component of the wave. It is wrong, however, to say that the electric component of an electromagnetic wave is “stronger” than the magnetic component. You cannot compare quantities that are measured in different units. However, these electric and magnetic components are on an equal basis because their average energies, which *can* be compared, are equal.

33-6 Radiation Pressure

Electromagnetic waves have linear momentum and thus can exert a pressure on an object when shining on it. However, the pressure must be very small because, for example, you do not feel a punch during a camera flash.

To find an expression for the pressure, let us shine a beam of electromagnetic radiation—light, for example—on an object for a time interval Δt . Further, let us assume that the object is free to move and that the radiation is entirely **absorbed** (taken up) by the object. This means that during the interval Δt , the object gains an energy ΔU from the radiation. Maxwell showed that the object also gains linear momentum. The magnitude Δp of the momentum change of the object is related to the energy change ΔU by

$$\Delta p = \frac{\Delta U}{c} \quad (\text{total absorption}), \quad (33-28)$$

where c is the speed of light. The direction of the momentum change of the object is the direction of the *incident* (incoming) beam that the object absorbs.

Instead of being absorbed, the radiation can be **reflected** by the object; that is, the radiation can be sent off in a new direction as if it bounced off the object. If the radiation is entirely reflected back along its original path, the magnitude of the momentum change of the object is twice that given above, or

$$\Delta p = \frac{2 \Delta U}{c} \quad (\text{total reflection back along path}). \quad (33-29)$$

In the same way, an object undergoes twice as much momentum change when a perfectly elastic tennis ball is bounced from it as when it is struck by a perfectly inelastic ball (a lump of wet putty, say) of the same mass and velocity. If the incident radiation is partly absorbed and partly reflected, the momentum change of the object is between $\Delta U/c$ and $2 \Delta U/c$.

From Newton's second law in its linear momentum form (Section 9-4), we know that a change in momentum is related to a force by

$$F = \frac{\Delta p}{\Delta t}. \quad (33-30)$$

To find expressions for the force exerted by radiation in terms of the intensity I of the radiation, we first note that intensity is

$$I = \frac{\text{power}}{\text{area}} = \frac{\text{energy/time}}{\text{area}}.$$

Next, suppose that a flat surface of area A , perpendicular to the path of the radiation, intercepts the radiation. In time interval Δt , the energy intercepted by area A is

$$\Delta U = IA \Delta t. \quad (33-31)$$

If the energy is completely absorbed, then Eq. 33-28 tells us that $\Delta p = IA \Delta t/c$, and, from Eq. 33-30, the magnitude of the force on the area A is

$$F = \frac{IA}{c} \quad (\text{total absorption}). \quad (33-32)$$

Similarly, if the radiation is totally reflected back along its original path, Eq. 33-29 tells us that $\Delta p = 2IA \Delta t/c$ and, from Eq. 33-30,

$$F = \frac{2IA}{c} \quad (\text{total reflection back along path}). \quad (33-33)$$

If the radiation is partly absorbed and partly reflected, the magnitude of the force on area A is between the values of IA/c and $2IA/c$.

The force per unit area on an object due to radiation is the radiation pressure p_r . We can find it for the situations of Eqs. 33-32 and 33-33 by dividing both sides of each equation by A . We obtain

$$p_r = \frac{I}{c} \quad (\text{total absorption}) \quad (33-34)$$

$$\text{and} \quad p_r = \frac{2I}{c} \quad (\text{total reflection back along path}). \quad (33-35)$$

Be careful not to confuse the symbol p_r for radiation pressure with the symbol p for momentum. Just as with fluid pressure in Chapter 14, the SI unit of radiation pressure is the newton per square meter (N/m^2), which is called the pascal (Pa).

The development of laser technology has permitted researchers to achieve radiation pressures much greater than, say, that due to a camera flashlamp. This comes about because a beam of laser light—unlike a beam of light from a small lamp filament—can be focused to a tiny spot. This permits the delivery of great amounts of energy to small objects placed at that spot.

CHECKPOINT 3

Light of uniform intensity shines perpendicularly on a totally absorbing surface, fully illuminating the surface. If the area of the surface is decreased, do (a) the radiation pressure and (b) the radiation force on the surface increase, decrease, or stay the same?

33-7 POLARIZATION

901

33-7 Polarization

VHF (very high frequency) television antennas in England are oriented vertically, but those in North America are horizontal. The difference is due to the direction of oscillation of the electromagnetic waves carrying the TV signal. In England, the transmitting equipment is designed to produce waves that are **polarized** vertically; that is, their electric field oscillates vertically. Thus, for the electric field of the incident television waves to drive a current along an antenna (and provide a signal to a television set), the antenna must be vertical. In North America, the waves are polarized horizontally.

Figure 33-9a shows an electromagnetic wave with its electric field oscillating parallel to the vertical y axis. The plane containing the \vec{E} vectors is called the **plane of oscillation** of the wave (hence, the wave is said to be *plane-polarized* in the y direction). We can represent the wave's *polarization* (state of being polarized) by showing the directions of the electric field oscillations in a head-on view of the plane of oscillation, as in Fig. 33-9b. The vertical double arrow in that figure indicates that as the wave travels past us, its electric field oscillates vertically—it continuously changes between being directed up and down the y axis.

Polarized Light

The electromagnetic waves emitted by a television station all have the same polarization, but the electromagnetic waves emitted by any common source of light (such as the Sun or a bulb) are **polarized randomly**, or **unpolarized** (the two terms mean the same thing). That is, the electric field at any given point is always perpendicular to the direction of travel of the waves but changes directions randomly. Thus, if we try to represent a head-on view of the oscillations over some time period, we do not have a simple drawing with a single double arrow like that of Fig. 33-9b; instead we have a mess of double arrows like that in Fig. 33-10a.

In principle, we can simplify the mess by resolving each electric field of Fig. 33-10a into y and z components. Then as the wave travels past us, the net y component oscillates parallel to the y axis and the net z component oscillates parallel to the z axis. We can then represent the unpolarized light with a pair of double arrows as shown in Fig. 33-10b. The double arrow along the y axis represents the oscillations of the net y component of the electric field. The double arrow along the z axis represents the oscillations of the net z component of the electric field. In doing all this, we effectively change unpolarized light into the superposition of two polarized waves whose planes of oscillation are perpendicular to each other—one plane contains the y axis and the other contains the z axis. One reason to make this change is that drawing Fig. 33-10b is a lot easier than drawing Fig. 33-10a.

We can draw similar figures to represent light that is **partially polarized** (its field oscillations are not completely random as in Fig. 33-10a, nor are they parallel to a single axis as in Fig. 33-9b). For this situation, we draw one of the double arrows in a perpendicular pair of double arrows longer than the other one.

Fig. 33-10 (a) Unpolarized light consists of waves with randomly directed electric fields. Here the waves are all traveling along the same axis, directly out of the page, and all have the same amplitude E . (b) A second way of representing unpolarized light—the light is the superposition of two polarized waves whose planes of oscillation are perpendicular to each other.

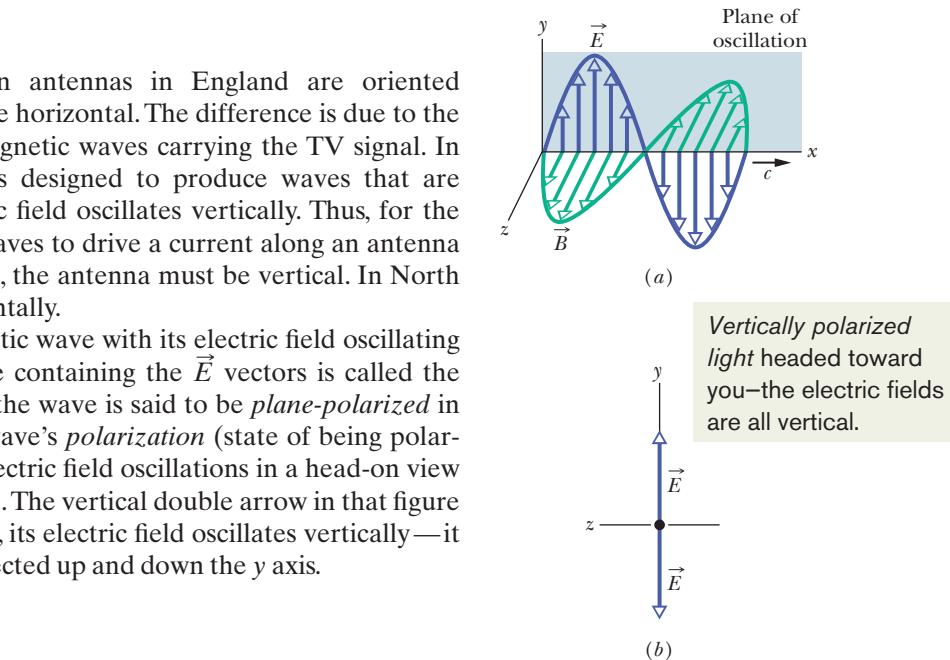
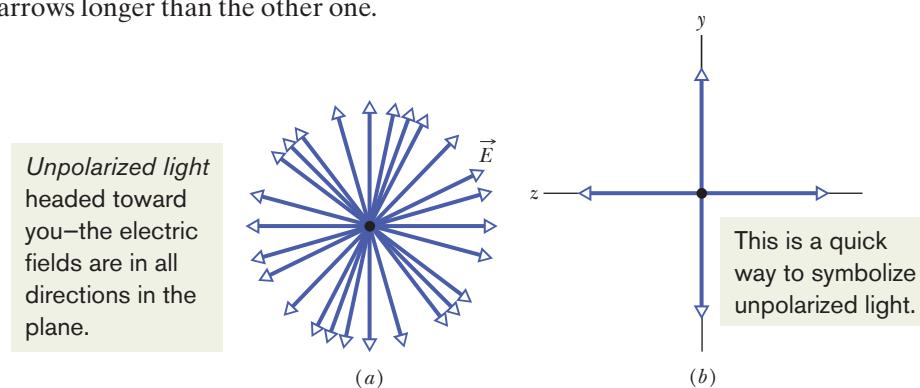


Fig. 33-9 (a) The plane of oscillation of a polarized electromagnetic wave. (b) To represent the polarization, we view the plane of oscillation head-on and indicate the directions of the oscillating electric field with a double arrow.



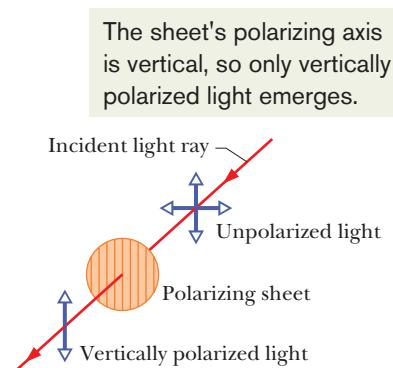


Fig. 33-11 Unpolarized light becomes polarized when it is sent through a polarizing sheet. Its direction of polarization is then parallel to the polarizing direction of the sheet, which is represented here by the vertical lines drawn in the sheet.

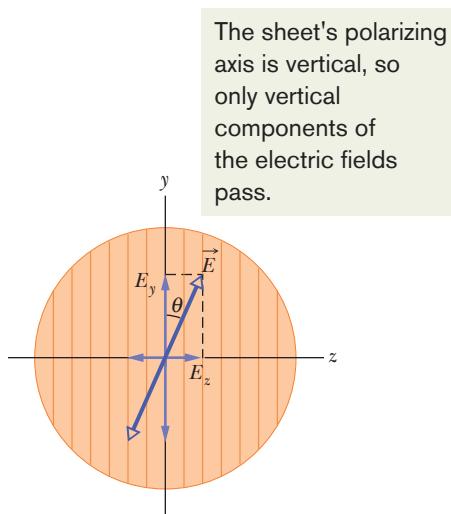


Fig. 33-12 Polarized light approaching a polarizing sheet. The electric field \vec{E} of the light can be resolved into components E_y (parallel to the polarizing direction of the sheet) and E_z (perpendicular to that direction). Component E_y will be transmitted by the sheet; component E_z will be absorbed.

We can transform unpolarized visible light into polarized light by sending it through a *polarizing sheet*, as is shown in Fig. 33-11. Such sheets, commercially known as Polaroids or Polaroid filters, were invented in 1932 by Edwin Land while he was an undergraduate student. A polarizing sheet consists of certain long molecules embedded in plastic. When the sheet is manufactured, it is stretched to align the molecules in parallel rows, like rows in a plowed field. When light is then sent through the sheet, electric field components along one direction pass through the sheet, while components perpendicular to that direction are absorbed by the molecules and disappear.

We shall not dwell on the molecules but, instead, shall assign to the sheet a *polarizing direction*, along which electric field components are passed:



An electric field component parallel to the polarizing direction is passed (*transmitted*) by a polarizing sheet; a component perpendicular to it is absorbed.

Thus, the electric field of the light emerging from the sheet consists of only the components that are parallel to the polarizing direction of the sheet; hence the light is polarized in that direction. In Fig. 33-11, the vertical electric field components are transmitted by the sheet; the horizontal components are absorbed. The transmitted waves are then vertically polarized.

Intensity of Transmitted Polarized Light

We now consider the intensity of light transmitted by a polarizing sheet. We start with unpolarized light, whose electric field oscillations we can resolve into y and z components as represented in Fig. 33-10b. Further, we can arrange for the y axis to be parallel to the polarizing direction of the sheet. Then only the y components of the light's electric field are passed by the sheet; the z components are absorbed. As suggested by Fig. 33-10b, if the original waves are randomly oriented, the sum of the y components and the sum of the z components are equal. When the z components are absorbed, half the intensity I_0 of the original light is lost. The intensity I of the emerging polarized light is then

$$I = \frac{1}{2}I_0. \quad (33-36)$$

Let us call this the *one-half rule*; we can use it *only* when the light reaching a polarizing sheet is unpolarized.

Suppose now that the light reaching a polarizing sheet is already polarized. Figure 33-12 shows a polarizing sheet in the plane of the page and the electric field \vec{E} of such a polarized light wave traveling toward the sheet (and thus prior to any absorption). We can resolve \vec{E} into two components relative to the polarizing direction of the sheet: parallel component E_y is transmitted by the sheet, and perpendicular component E_z is absorbed. Since θ is the angle between \vec{E} and the polarizing direction of the sheet, the transmitted parallel component is

$$E_y = E \cos \theta. \quad (33-37)$$

Recall that the intensity of an electromagnetic wave (such as our light wave) is proportional to the square of the electric field's magnitude (Eq. 33-26, $I = E_{\text{rms}}^2/c\mu_0$). In our present case then, the intensity I of the emerging wave is proportional to E_y^2 and the intensity I_0 of the original wave is proportional to E^2 . Hence, from Eq. 33-37 we can write $I/I_0 = \cos^2 \theta$, or

$$I = I_0 \cos^2 \theta. \quad (33-38)$$

Let us call this the *cosine-squared rule*; we can use it *only* when the light reaching

33-7 POLARIZATION

903

a polarizing sheet is already polarized. Then the transmitted intensity I is a maximum and is equal to the original intensity I_0 when the original wave is polarized parallel to the polarizing direction of the sheet (when θ in Eq. 33-38 is 0° or 180°). The transmitted intensity is zero when the original wave is polarized perpendicular to the polarizing direction of the sheet (when θ is 90°).

Figure 33-13 shows an arrangement in which initially unpolarized light is sent through two polarizing sheets P_1 and P_2 . (Often, the first sheet is called the *polarizer*, and the second the *analyzer*.) Because the polarizing direction of P_1 is vertical, the light transmitted by P_1 to P_2 is polarized vertically. If the polarizing direction of P_2 is also vertical, then all the light transmitted by P_1 is transmitted by P_2 . If the polarizing direction of P_2 is horizontal, none of the light transmitted by P_1 is transmitted by P_2 . We reach the same conclusions by considering only the *relative* orientations of the two sheets: If their polarizing directions are parallel, all the light passed by the first sheet is passed by the second sheet (Fig. 33-14a). If those directions are perpendicular (the sheets are said to be *crossed*), no light is passed by the second sheet (Fig. 33-14b). Finally, if the two polarizing directions of Fig. 33-13 make an angle between 0° and 90° , some of the light transmitted by P_1 will be transmitted by P_2 , as set by Eq. 33-38.

Light can be polarized by means other than polarizing sheets, such as by reflection (discussed in Section 33-10) and by scattering from atoms or molecules. In *scattering*, light that is intercepted by an object, such as a molecule, is sent off in many, perhaps random, directions. An example is the scattering of sunlight by molecules in the atmosphere, which gives the sky its general glow.

Although direct sunlight is unpolarized, light from much of the sky is at least partially polarized by such scattering. Bees use the polarization of sky light in navigating to and from their hives. Similarly, the Vikings used it to navigate across the North Sea when the daytime Sun was below the horizon (because of the high latitude of the North Sea). These early seafarers had discovered certain crystals (now called cordierite) that changed color when rotated in polarized light. By looking at the sky through such a crystal while rotating it about their line of sight, they could locate the hidden Sun and thus determine which way was south.

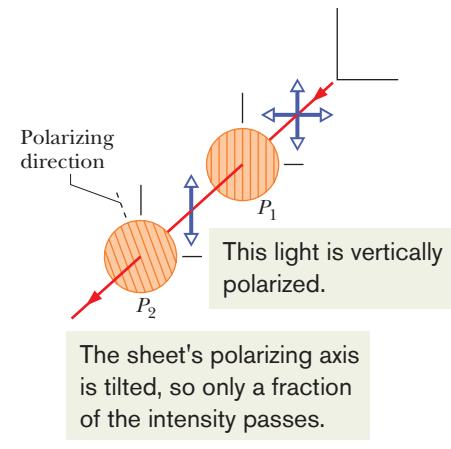


Fig. 33-13 The light transmitted by polarizing sheet P_1 is vertically polarized, as represented by the vertical double arrow. The amount of that light that is then transmitted by polarizing sheet P_2 depends on the angle between the polarization direction of that light and the polarizing direction of P_2 (indicated by the lines drawn in the sheet and by the dashed line).

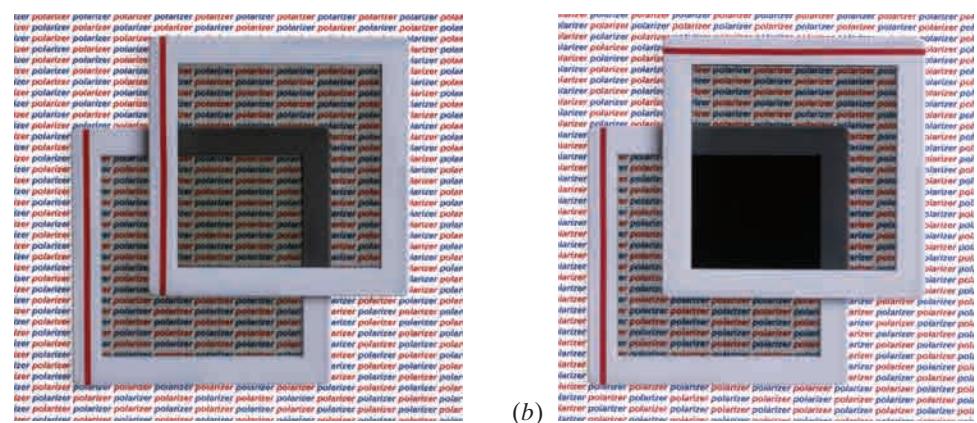


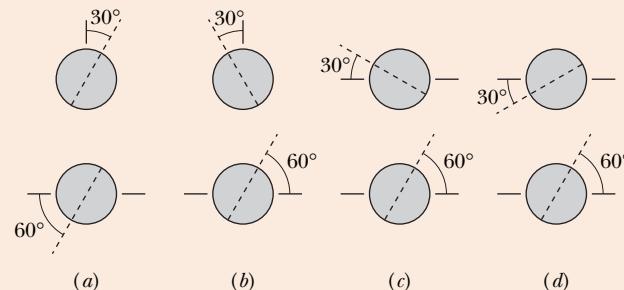
Fig. 33-14 (a) Overlapping polarizing sheets transmit light fairly well when their polarizing directions have the same orientation, but (b) they block most of the light when they are crossed. (Richard Megna/Fundamental Photographs.)

(a)

(b)

**CHECKPOINT 4**

The figure shows four pairs of polarizing sheets, seen face-on. Each pair is mounted in the path of initially unpolarized light. The polarizing direction of each sheet (indicated by the dashed line) is referenced to either a horizontal x axis or a vertical y axis. Rank the pairs according to the fraction of the initial intensity that they pass, greatest first.



Sample Problem

Polarization and intensity with three polarizing sheets

Figure 33-15a, drawn in perspective, shows a system of three polarizing sheets in the path of initially unpolarized light. The polarizing direction of the first sheet is parallel to the y axis, that of the second sheet is at an angle of 60° counterclockwise from the y axis, and that of the third sheet is parallel to the x axis. What fraction of the initial intensity I_0 of the light emerges from the three-sheet system, and in which direction is that emerging light polarized?

KEY IDEAS

1. We work through the system sheet by sheet, from the first one encountered by the light to the last one.
2. To find the intensity transmitted by any sheet, we apply either the one-half rule or the cosine-squared rule, depending on whether the light reaching the sheet is unpolarized or already polarized.
3. The light that is transmitted by a polarizing sheet is always polarized parallel to the polarizing direction of the sheet.

First sheet: The original light wave is represented in Fig. 33-15b, using the head-on, double-arrow representation of Fig. 33-10b. Because the light is initially unpolarized, the intensity I_1 of the light transmitted by the first sheet is given by the one-half rule (Eq. 33-36):

$$I_1 = \frac{1}{2} I_0.$$

Because the polarizing direction of the first sheet is parallel to the y axis, the polarization of the light transmitted by it is also, as shown in the head-on view of Fig. 33-15c.

Second sheet: Because the light reaching the second sheet is polarized, the intensity I_2 of the light transmitted by that sheet is given by the cosine-squared rule (Eq. 33-38). The angle

θ in the rule is the angle between the polarization direction of the entering light (parallel to the y axis) and the polarizing direction of the second sheet (60° counterclockwise from the y axis), and so θ is 60° . (The larger angle between the two directions, namely 120° , can also be used.) We have

$$I_2 = I_1 \cos^2 60^\circ.$$

The polarization of this transmitted light is parallel to the polarizing direction of the sheet transmitting it—that is, 60° counterclockwise from the y axis, as shown in the head-on view of Fig. 33-15d.

Third sheet: Because the light reaching the third sheet is polarized, the intensity I_3 of the light transmitted by that sheet is given by the cosine-squared rule. The angle θ is now the angle between the polarization direction of the entering light (Fig. 33-15d) and the polarizing direction of the third sheet (parallel to the x axis), and so $\theta = 30^\circ$. Thus,

$$I_3 = I_2 \cos^2 30^\circ.$$

This final transmitted light is polarized parallel to the x axis (Fig. 33-15e). We find its intensity by substituting first for I_2 and then for I_1 in the equation above:

$$\begin{aligned} I_3 &= I_2 \cos^2 30^\circ = (I_1 \cos^2 60^\circ) \cos^2 30^\circ \\ &= (\frac{1}{2} I_0) \cos^2 60^\circ \cos^2 30^\circ = 0.094 I_0. \end{aligned}$$

Thus,
$$\frac{I_3}{I_0} = 0.094. \quad (\text{Answer})$$

That is to say, 9.4% of the initial intensity emerges from the three-sheet system. (If we now remove the second sheet, what fraction of the initial intensity emerges from the system?)



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33-8 Reflection and Refraction

Although a light wave spreads as it moves away from its source, we can often approximate its travel as being in a straight line; we did so for the light wave in Fig. 33-5a. The study of the properties of light waves under that approximation is called *geometrical optics*. For the rest of this chapter and all of Chapter 34, we shall discuss the geometrical optics of visible light.

The photograph in Fig. 33-16a shows an example of light waves traveling in approximately straight lines. A narrow beam of light (the *incident beam*), angled downward from the left and traveling through air, encounters a *plane* (flat) water surface. Part of the light is **reflected** by the surface, forming a beam directed upward toward the right, traveling as if the original beam had bounced from the surface. The rest of the light travels through the surface and into the water, form-

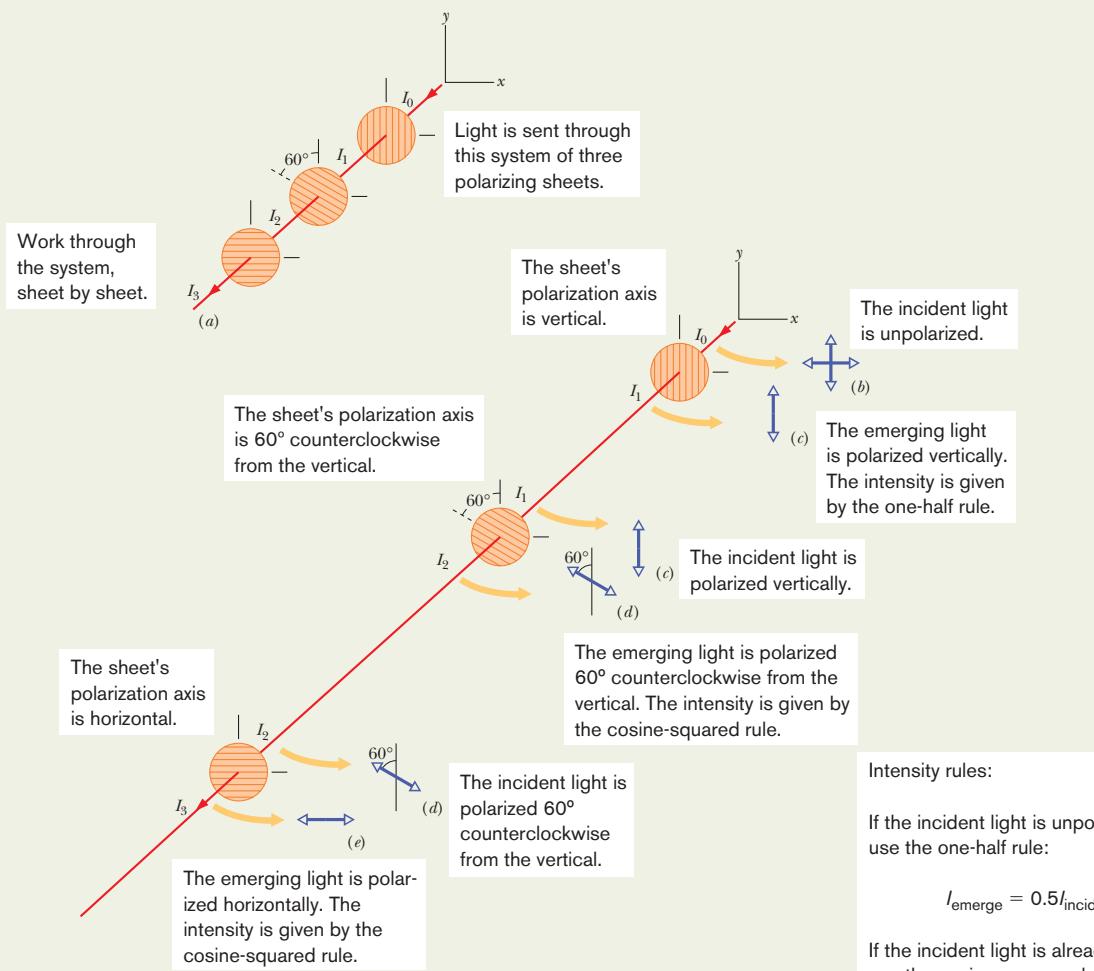


Fig. 33-15 (a) Initially unpolarized light of intensity I_0 is sent into a system of three polarizing sheets. The intensities I_1 , I_2 , and I_3 of the light transmitted by the sheets are labeled. Shown also are the polarizations, from head-on views, of (b) the initial light and the light transmitted by (c) the first sheet, (d) the second sheet, and (e) the third sheet.

Intensity rules:

If the incident light is unpolarized, use the one-half rule:

$$I_{\text{emerge}} = 0.5 I_{\text{incident}}$$

If the incident light is already polarized, use the cosine-square rule:

$$I_{\text{emerge}} = I_{\text{incident}} (\cos \theta)^2,$$

but be sure to insert the angle between the polarization of the incident light and the polarization axis of the sheet.

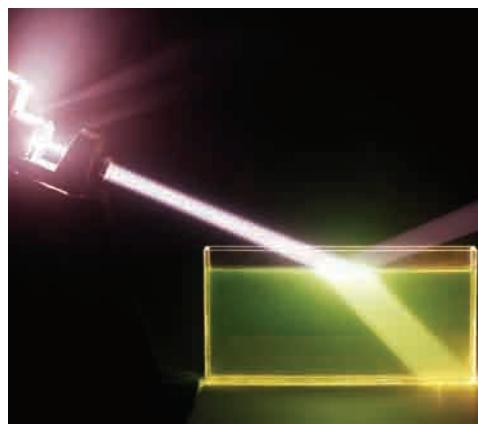


Fig. 33-16 (a) A photograph showing an incident beam of light reflected and refracted by a horizontal water surface. (©1974 FP/Fundamentals Photography) (Fig. 33-16 continues on next page.)

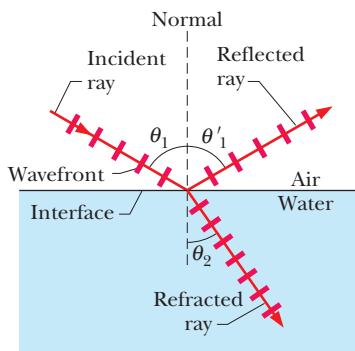


Fig. 33-16 (Continued) (b) A ray representation of (a). The angles of incidence (θ_1), reflection (θ'_1), and refraction (θ_2) are marked.

ing a beam directed downward to the right. Because light can travel through it, the water is said to be *transparent*; that is, we can see through it. (In this chapter we shall consider only transparent materials and not opaque materials, through which light cannot travel.)

The travel of light through a surface (or *interface*) that separates two media is called **refraction**, and the light is said to be *refracted*. Unless an incident beam of light is perpendicular to the surface, refraction changes the light's direction of travel. For this reason, the beam is said to be "bent" by the refraction. Note in Fig. 33-16a that the bending occurs only at the surface; within the water, the light travels in a straight line.

In Figure 33-16b, the beams of light in the photograph are represented with an *incident ray*, a *reflected ray*, and a *refracted ray* (and wavefronts). Each ray is oriented with respect to a line, called the *normal*, that is perpendicular to the surface at the point of reflection and refraction. In Fig. 33-16b, the **angle of incidence** is θ_1 , the **angle of reflection** is θ'_1 , and the **angle of refraction** is θ_2 , all measured *relative to the normal*. The plane containing the incident ray and the normal is the *plane of incidence*, which is in the plane of the page in Fig. 33-16b.

Experiment shows that reflection and refraction are governed by two laws:

Law of reflection: A reflected ray lies in the plane of incidence and has an angle of reflection equal to the angle of incidence (both relative to the normal). In Fig. 33-16b, this means that

$$\theta'_1 = \theta_1 \quad (\text{reflection}). \quad (33-39)$$

(We shall now usually drop the prime on the angle of reflection.)

Law of refraction: A refracted ray lies in the plane of incidence and has an angle of refraction θ_2 that is related to the angle of incidence θ_1 by

$$n_2 \sin \theta_2 = n_1 \sin \theta_1 \quad (\text{refraction}). \quad (33-40)$$

Here each of the symbols n_1 and n_2 is a dimensionless constant, called the **index of refraction**, that is associated with a medium involved in the refraction. We derive this equation, called **Snell's law**, in Chapter 35. As we shall discuss there, the index of refraction of a medium is equal to c/v , where v is the speed of light in that medium and c is its speed in vacuum.

Table 33-1 gives the indexes of refraction of vacuum and some common substances. For vacuum, n is defined to be exactly 1; for air, n is very close to 1.0 (an approximation we shall often make). Nothing has an index of refraction below 1.

Table 33-1

Some Indexes of Refraction^a

Medium	Index	Medium	Index
Vacuum	Exactly 1	Typical crown glass	1.52
Air (STP) ^b	1.00029	Sodium chloride	1.54
Water (20°C)	1.33	Polystyrene	1.55
Acetone	1.36	Carbon disulfide	1.63
Ethyl alcohol	1.36	Heavy flint glass	1.65
Sugar solution (30%)	1.38	Sapphire	1.77
Fused quartz	1.46	Heaviest flint glass	1.89
Sugar solution (80%)	1.49	Diamond	2.42

^aFor a wavelength of 589 nm (yellow sodium light).

^bSTP means "standard temperature (0°C) and pressure (1 atm)."

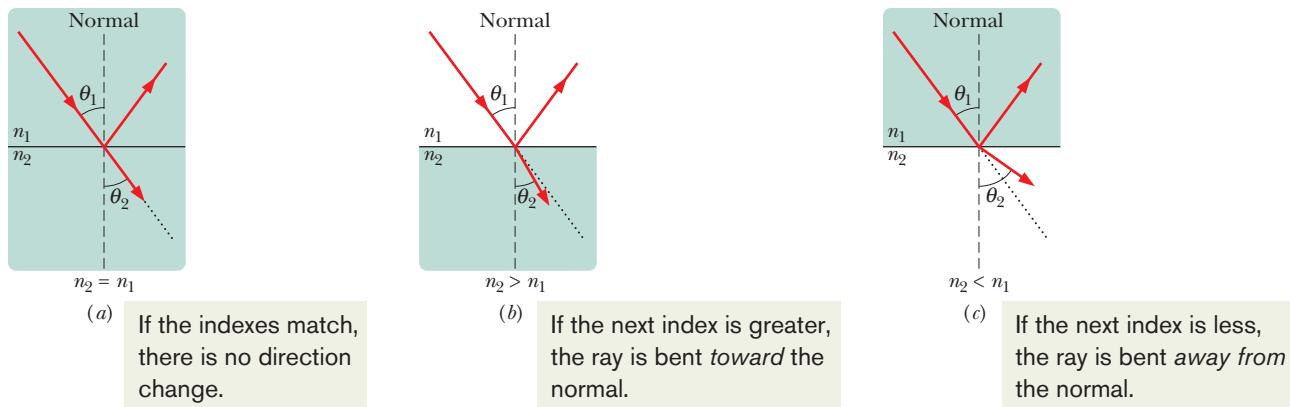


Fig. 33-17 Refraction of light traveling from a medium with an index of refraction n_1 into a medium with an index of refraction n_2 . (a) The beam does not bend when $n_2 = n_1$; the refracted light then travels in the *undeflected direction* (the dotted line), which is the same as the direction of the incident beam. The beam bends (b) toward the normal when $n_2 > n_1$ and (c) away from the normal when $n_2 < n_1$.

We can rearrange Eq. 33-40 as

$$\sin \theta_2 = \frac{n_1}{n_2} \sin \theta_1 \quad (33-41)$$

to compare the angle of refraction θ_2 with the angle of incidence θ_1 . We can then see that the relative value of θ_2 depends on the relative values of n_2 and n_1 :

1. If n_2 is equal to n_1 , then θ_2 is equal to θ_1 and refraction does not bend the light beam, which continues in the *undeflected direction*, as in Fig. 33-17a.
2. If n_2 is greater than n_1 , then θ_2 is less than θ_1 . In this case, refraction bends the light beam away from the undeflected direction and toward the normal, as in Fig. 33-17b.
3. If n_2 is less than n_1 , then θ_2 is greater than θ_1 . In this case, refraction bends the light beam away from the undeflected direction and away from the normal, as in Fig. 33-17c.

Refraction *cannot* bend a beam so much that the refracted ray is on the same side of the normal as the incident ray.

Chromatic Dispersion

The index of refraction n encountered by light in any medium except vacuum depends on the wavelength of the light. The dependence of n on wavelength implies that when a light beam consists of rays of different wavelengths, the rays will be refracted at different angles by a surface; that is, the light will be spread out by the refraction. This spreading of light is called **chromatic dispersion**, in which “chromatic” refers to the colors associated with the individual wavelengths and “dispersion” refers to the spreading of the light according to its wavelengths or colors. The refractions of Figs. 33-16 and 33-17 do not show chromatic dispersion because the beams are *monochromatic* (of a single wavelength or color).

Generally, the index of refraction of a given medium is *greater* for a shorter wavelength (corresponding to, say, blue light) than for a longer wavelength (say, red light). As an example, Fig. 33-18 shows how the index of refraction of fused quartz depends on the wavelength of light. Such dependence means that when a beam made up of waves of both blue and red light is refracted through a surface, such as from air into quartz or vice versa, the blue *component* (the ray corresponding to the wave of blue light) bends more than the red component.

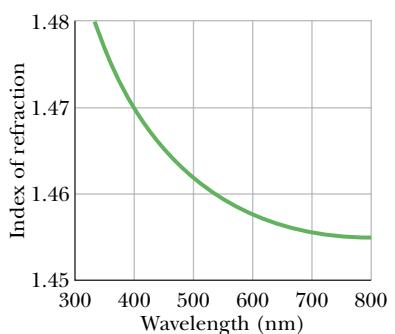


Fig. 33-18 The index of refraction as a function of wavelength for fused quartz. The graph indicates that a beam of short-wavelength light, for which the index of refraction is higher, is bent more upon entering or leaving quartz than a beam of long-wavelength light.

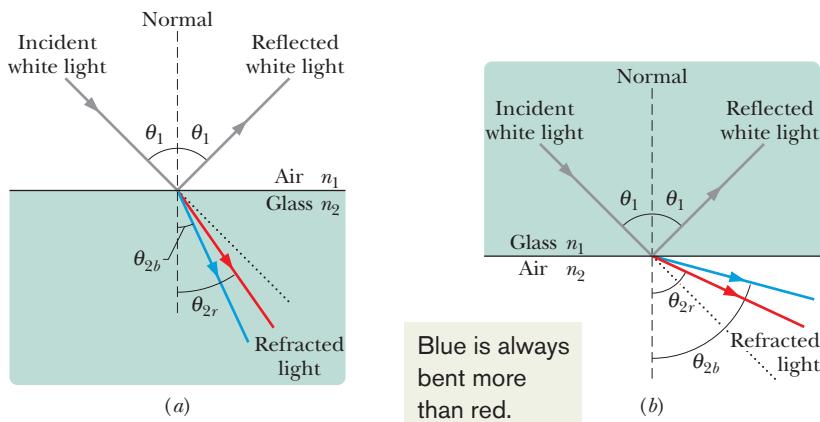


Fig. 33-19 Chromatic dispersion of white light. The blue component is bent more than the red component. (a) Passing from air to glass, the blue component ends up with the smaller angle of refraction. (b) Passing from glass to air, the blue component ends up with the greater angle of refraction. Each dotted line represents the direction in which the light would continue to travel if it were not bent by the refraction.

A beam of *white light* consists of components of all (or nearly all) the colors in the visible spectrum with approximately uniform intensities. When you see such a beam, you perceive white rather than the individual colors. In Fig. 33-19a, a beam of white light in air is incident on a glass surface. (Because the pages of this book are white, a beam of white light is represented with a gray ray here. Also, a beam of monochromatic light is generally represented with a red ray.) Of the refracted light in Fig. 33-19a, only the red and blue components are shown. Because the blue component is bent more than the red component, the angle of refraction θ_{2b} for the blue component is *smaller* than the angle of refraction θ_{2r} for the red component. (Remember, angles are measured relative to the normal.) In Fig. 33-19b, a ray of white light in glass is incident on a glass-air interface. Again, the blue component is bent more than the red component, but now θ_{2b} is greater than θ_{2r} .

To increase the color separation, we can use a solid glass prism with a triangular cross section, as in Fig. 33-20a. The dispersion at the first surface (on the left in Figs. 33-20a, b) is then enhanced by the dispersion at the second surface.

Rainbows

The most charming example of chromatic dispersion is a rainbow. When sunlight (which consists of all visible colors) is intercepted by a falling raindrop,

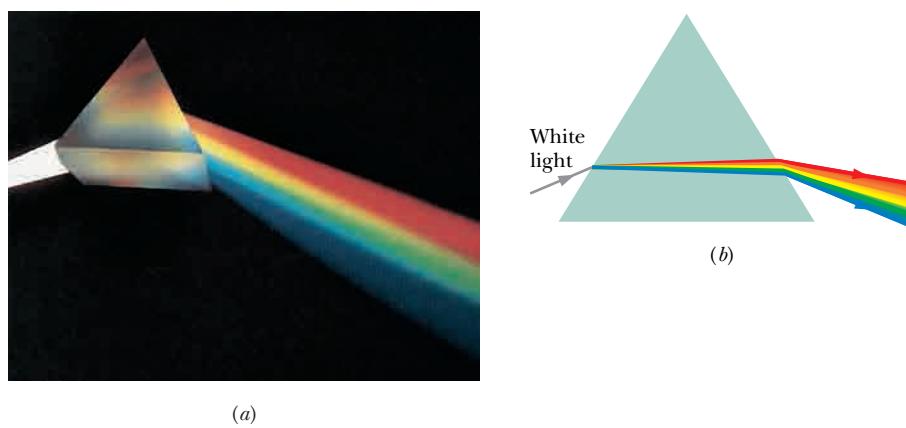


Fig. 33-20 (a) A triangular prism separating white light into its component colors. (b) Chromatic dispersion occurs at the first surface and is increased at the second surface. (*Courtesy Bausch & Lomb*)

33-8 REFLECTION AND REFRACTION

909

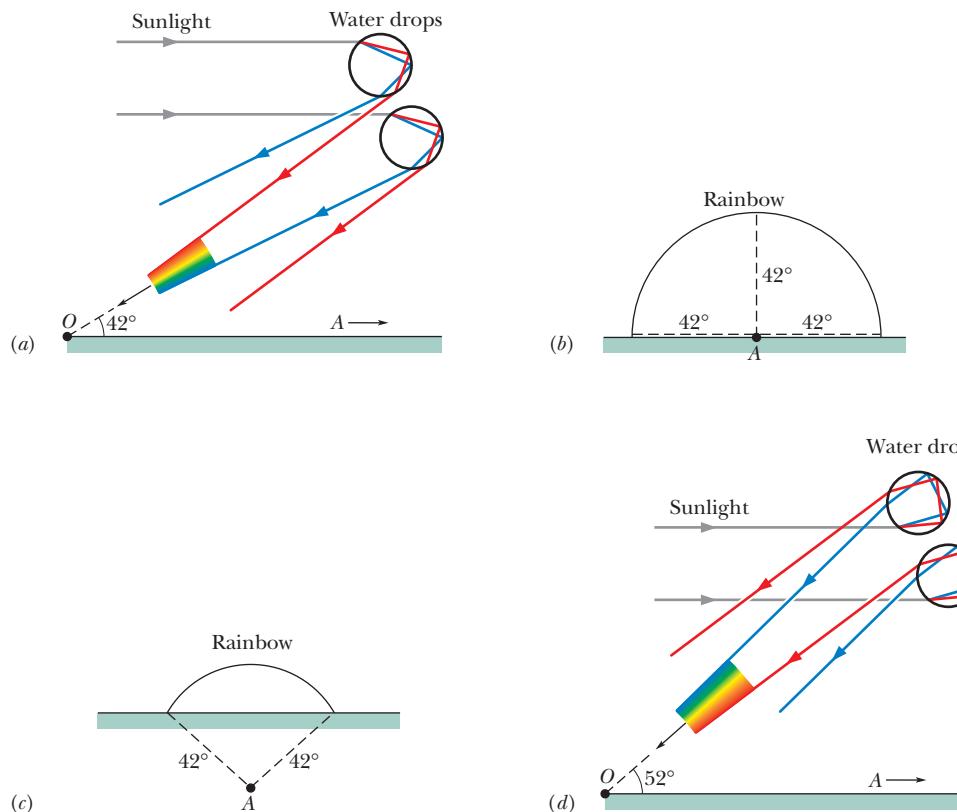


Fig. 33-21 (a) The separation of colors when sunlight refracts into and out of falling raindrops leads to a primary rainbow. The antisolar point A is on the horizon at the right. The rainbow colors appear at an angle of 42° from the direction of A . (b) Drops at 42° from A in any direction can contribute to the rainbow. (c) The rainbow arc when the Sun is higher (and thus A is lower). (d) The separation of colors leading to a secondary rainbow.

some of the light refracts into the drop, reflects once from the drop's inner surface, and then refracts out of the drop. Figure 33-21a shows the situation when the Sun is on the horizon at the left (and thus when the rays of sunlight are horizontal). The first refraction separates the sunlight into its component colors, and the second refraction increases the separation. (Only the red and blue rays are shown in the figure.) If many falling drops are brightly illuminated, you can see the separated colors they produce when the drops are at an angle of 42° from the direction of the *antisolar point* A , the point directly opposite the Sun in your view.

To locate the drops, face away from the Sun and point both arms directly away from the Sun, toward the shadow of your head. Then move your right arm directly up, directly rightward, or in any intermediate direction until the angle between your arms is 42° . If illuminated drops happen to be in the direction of your right arm, you see color in that direction.

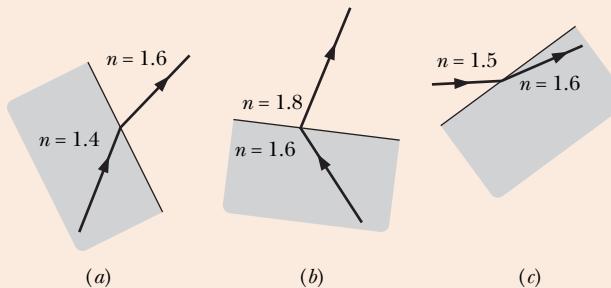
Because any drop at an angle of 42° in any direction from A can contribute to the rainbow, the rainbow is always a 42° circular arc around A (Fig. 33-21b) and the top of a rainbow is never more than 42° above the horizon. When the Sun is above the horizon, the direction of A is below the horizon, and only a shorter, lower rainbow arc is possible (Fig. 33-21c).

Because rainbows formed in this way involve one reflection of light inside each drop, they are often called *primary rainbows*. A *secondary rainbow* involves two reflections inside a drop, as shown in Fig. 33-21d. Colors appear in the secondary rainbow at an angle of 52° from the direction of A . A secondary rainbow is wider and dimmer than a primary rainbow and thus is more difficult to see. Also, the order of colors in a secondary rainbow is reversed from the order in a primary rainbow, as you can see by comparing parts a and d of Fig. 33-21.

Rainbows involving three or four reflections occur in the direction of the Sun and cannot be seen against the glare of sunshine in that part of the sky. Rainbows involving even more reflections inside the drops are too dim to see.

 **CHECKPOINT 5**

Which of the three drawings here (if any) show physically possible refraction?


Sample Problem
Reflection and refraction of a monochromatic beam

- (a) In Fig. 33-22a, a beam of monochromatic light reflects and refracts at point A on the interface between material 1 with index of refraction $n_1 = 1.33$ and material 2 with index of refraction $n_2 = 1.77$. The incident beam makes an angle of 50° with the interface. What is the angle of reflection at point A ? What is the angle of refraction there?

KEY IDEAS

- (1) The angle of reflection is equal to the angle of incidence, and both angles are measured relative to the normal to the surface at the point of reflection. (2) When light reaches the interface between two materials with different indexes of refraction (call them n_1 and n_2), part of the light can be refracted by the interface according to Snell's law, Eq. 33-40:

$$n_2 \sin \theta_2 = n_1 \sin \theta_1, \quad (33-42)$$

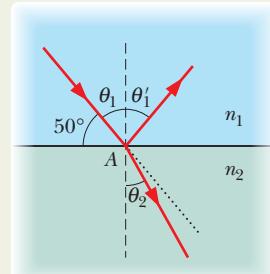
where both angles are measured relative to the normal at the point of refraction.

Calculations: In Fig. 33-22a, the normal at point A is drawn as a dashed line through the point. Note that the angle of incidence θ_1 is not the given 50° but is $90^\circ - 50^\circ = 40^\circ$. Thus, the angle of reflection is

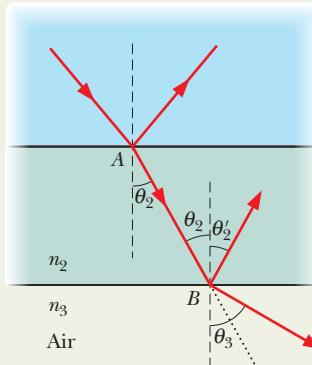
$$\theta'_1 = \theta_1 = 40^\circ. \quad (\text{Answer})$$

The light that passes from material 1 into material 2 undergoes refraction at point A on the interface between the two materials. Again we measure angles between light rays and a normal, here at the point of refraction. Thus, in Fig. 33-22a, the angle of refraction is the angle marked θ_2 . Solving Eq. 33-42 for θ_2 gives us

$$\begin{aligned} \theta_2 &= \sin^{-1} \left(\frac{n_1}{n_2} \sin \theta_1 \right) = \sin^{-1} \left(\frac{1.33}{1.77} \sin 40^\circ \right) \\ &= 28.88^\circ \approx 29^\circ. \end{aligned} \quad (\text{Answer})$$



(a)



(b)

Fig. 33-22 (a) Light reflects and refracts at point A on the interface between materials 1 and 2. (b) The light that passes through material 2 reflects and refracts at point B on the interface between materials 2 and 3 (air). Each dashed line is a normal. Each dotted line gives the incident direction of travel.

This result means that the beam swings toward the normal (it was at 40° to the normal and is now at 29°). The reason is that when the light travels across the interface, it moves into a material with a greater index of refraction. *Caution:* Note that the beam does *not* swing through the normal so that it appears on the left side of Fig. 33-22a.

33-9 TOTAL INTERNAL REFLECTION

911

(b) The light that enters material 2 at point *A* then reaches point *B* on the interface between material 2 and material 3, which is air, as shown in Fig. 33-22b. The interface through *B* is parallel to that through *A*. At *B*, some of the light reflects and the rest enters the air. What is the angle of reflection? What is the angle of refraction into the air?

Calculations: We first need to relate one of the angles at point *B* with a known angle at point *A*. Because the interface through point *B* is parallel to that through point *A*, the incident angle at *B* must be equal to the angle of refraction θ_2 , as shown in Fig. 33-22b. Then for reflection, we again use the law of reflection. Thus, the angle of reflection at *B* is

$$\theta'_2 = \theta_2 = 28.88^\circ \approx 29^\circ. \quad (\text{Answer})$$

Next, the light that passes from material 2 into the air undergoes refraction at point *B*, with refraction angle θ_3 . Thus, we again apply Snell's law of refraction, but this time we write Eq. 33-40 as

$$n_3 \sin \theta_3 = n_2 \sin \theta_2. \quad (33-43)$$

Solving for θ_3 then leads to

$$\begin{aligned} \theta_3 &= \sin^{-1} \left(\frac{n_2}{n_3} \sin \theta_2 \right) = \sin^{-1} \left(\frac{1.77}{1.00} \sin 28.88^\circ \right) \\ &= 58.75^\circ \approx 59^\circ. \end{aligned} \quad (\text{Answer})$$

This result means that the beam swings away from the normal (it was at 29° to the normal and is now at 59°). The reason is that when the light travels across the interface, it moves into a material (air) with a lower index of refraction.



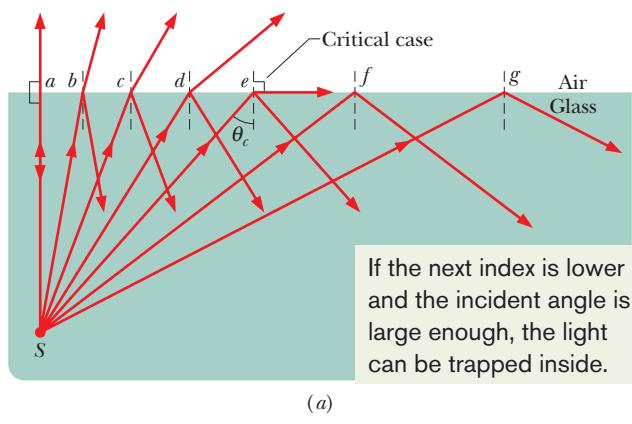
Additional examples, video, and practice available at WileyPLUS

33-9 Total Internal Reflection

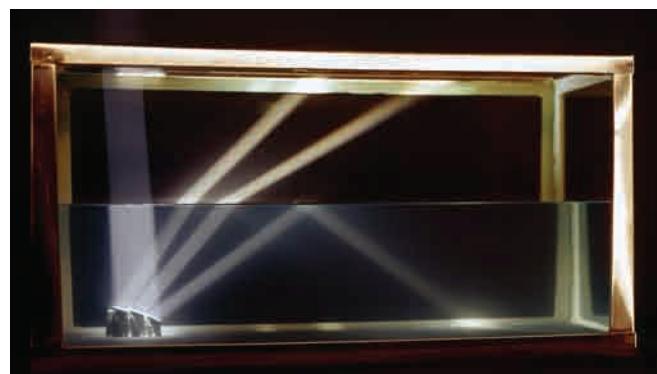
Figure 33-23a shows rays of monochromatic light from a point source *S* in glass incident on the interface between the glass and air. For ray *a*, which is perpendicular to the interface, part of the light reflects at the interface and the rest travels through it with no change in direction.

For rays *b* through *e*, which have progressively larger angles of incidence at the interface, there are also both reflection and refraction at the interface. As the angle of incidence increases, the angle of refraction increases; for ray *e* it is 90° , which means that the refracted ray points directly along the interface. The angle of incidence giving this situation is called the **critical angle** θ_c . For angles of incidence larger than θ_c , such as for rays *f* and *g*, there is no refracted ray and *all* the light is reflected; this effect is called **total internal reflection**.

To find θ_c , we use Eq. 33-40; we arbitrarily associate subscript 1 with the glass and subscript 2 with the air, and then we substitute θ_c for θ_1 and 90° for θ_2 ,



(a)

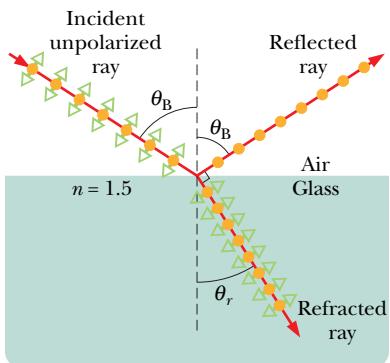


(b)

Fig. 33-23 (a) Total internal reflection of light from a point source *S* in glass occurs for all angles of incidence greater than the critical angle θ_c . At the critical angle, the refracted ray points along the air–glass interface. (b) A source in a tank of water. (Ken Kay/Fundamental Photographs)



Fig. 33-24 An endoscope used to inspect an artery. (©Laurent/Phototake)



- Component perpendicular to page
- ↔ Component parallel to page

Fig. 33-25 A ray of unpolarized light in air is incident on a glass surface at the Brewster angle θ_B . The electric fields along that ray have been resolved into components perpendicular to the page (the plane of incidence, reflection, and refraction) and components parallel to the page. The reflected light consists only of components perpendicular to the page and is thus polarized in that direction. The refracted light consists of the original components parallel to the page and weaker components perpendicular to the page; this light is partially polarized.

which leads to

$$n_1 \sin \theta_c = n_2 \sin 90^\circ, \quad (33-44)$$

which gives us

$$\theta_c = \sin^{-1} \frac{n_2}{n_1} \quad (\text{critical angle}). \quad (33-45)$$

Because the sine of an angle cannot exceed unity, n_2 cannot exceed n_1 in this equation. This restriction tells us that total internal reflection cannot occur when the incident light is in the medium of lower index of refraction. If source S were in the air in Fig. 33-23a, all its rays that are incident on the air–glass interface (including f and g) would be both reflected *and* refracted at the interface.

Total internal reflection has found many applications in medical technology. For example, a physician can view the interior of an artery of a patient by running two thin bundles of *optical fibers* through the chest wall and into an artery (Fig. 33-24). Light introduced at the outer end of one bundle undergoes repeated total internal reflection within the fibers so that, even though the bundle provides a curved path, most of the light ends up exiting the other end and illuminating the interior of the artery. Some of the light reflected from the interior then comes back up the second bundle in a similar way, to be detected and converted to an image on a monitor's screen for the physician to view.

33-10 Polarization by Reflection

You can vary the glare you see in sunlight that has been reflected from, say, water by looking through a polarizing sheet (such as a polarizing sunglass lens) and then rotating the sheet's polarizing axis around your line of sight. You can do so because any light that is reflected from a surface is either fully or partially polarized by the reflection.

Figure 33-25 shows a ray of unpolarized light incident on a glass surface. Let us resolve the electric field vectors of the light into two components. The *perpendicular components* are perpendicular to the plane of incidence and thus also to the page in Fig. 33-25; these components are represented with dots (as if we see the tips of the vectors). The *parallel components* are parallel to the plane of incidence and the page; they are represented with double-headed arrows. Because the light is unpolarized, these two components are of equal magnitude.

In general, the reflected light also has both components but with unequal magnitudes. This means that the reflected light is partially polarized—the electric fields oscillating along one direction have greater amplitudes than those oscillating along other directions. However, when the light is incident at a particular incident angle, called the *Brewster angle* θ_B , the reflected light has only perpendicular components, as shown in Fig. 33-25. The reflected light is then fully polarized perpendicular to the plane of incidence. The parallel components of the incident light do not disappear but (along with perpendicular components) refract into the glass.

Glass, water, and the other dielectric materials discussed in Section 25-7 can partially and fully polarize light by reflection. When you intercept sunlight reflected from such a surface, you see a bright spot (the glare) on the surface where the reflection takes place. If the surface is horizontal as in Fig. 33-25, the reflected light is partially or fully polarized horizontally. To eliminate such glare from horizontal surfaces, the lenses in polarizing sunglasses are mounted with their polarizing direction vertical.

Brewster's Law

For light incident at the Brewster angle θ_B , we find experimentally that the reflected and refracted rays are perpendicular to each other. Because the reflected ray is reflected at the angle θ_B in Fig. 33-25 and the refracted ray is at an angle θ_r , we have

$$\theta_B + \theta_r = 90^\circ. \quad (33-46)$$

These two angles can also be related with Eq. 33-40. Arbitrarily assigning subscript 1 in Eq. 33-40 to the material through which the incident and reflected rays travel, we have, from that equation,

$$n_1 \sin \theta_B = n_2 \sin \theta_r. \quad (33-47)$$

Combining these equations leads to

$$n_1 \sin \theta_B = n_2 \sin(90^\circ - \theta_B) = n_2 \cos \theta_B, \quad (33-48)$$

which gives us

$$\theta_B = \tan^{-1} \frac{n_2}{n_1} \quad (\text{Brewster angle}). \quad (33-49)$$

(Note carefully that the subscripts in Eq. 33-49 are *not* arbitrary because of our decision as to their meanings.) If the incident and reflected rays travel *in air*, we can approximate n_1 as unity and let n represent n_2 in order to write Eq. 33-49 as

$$\theta_B = \tan^{-1} n \quad (\text{Brewster's law}). \quad (33-50)$$

This simplified version of Eq. 33-49 is known as **Brewster's law**. Like θ_B , it is named after Sir David Brewster, who found both experimentally in 1812.

REVIEW & SUMMARY

Electromagnetic Waves An electromagnetic wave consists of oscillating electric and magnetic fields. The various possible frequencies of electromagnetic waves form a *spectrum*, a small part of which is visible light. An electromagnetic wave traveling along an x axis has an electric field \vec{E} and a magnetic field \vec{B} with magnitudes that depend on x and t :

$$E = E_m \sin(kx - \omega t)$$

and

$$B = B_m \sin(kx - \omega t), \quad (33-1, 33-2)$$

where E_m and B_m are the amplitudes of \vec{E} and \vec{B} . The electric field induces the magnetic field and vice versa. The speed of any electromagnetic wave in vacuum is c , which can be written as

$$c = \frac{E}{B} = \frac{1}{\sqrt{\mu_0 \epsilon_0}}, \quad (33-5, 33-3)$$

where E and B are the simultaneous magnitudes of the fields.

Energy Flow The rate per unit area at which energy is transported via an electromagnetic wave is given by the Poynting vector \vec{S} :

$$\vec{S} = \frac{1}{\mu_0} \vec{E} \times \vec{B}. \quad (33-19)$$

The direction of \vec{S} (and thus of the wave's travel and the energy transport) is perpendicular to the directions of both \vec{E} and \vec{B} . The time-averaged rate per unit area at which energy is transported is

S_{avg} , which is called the *intensity* I of the wave:

$$I = \frac{1}{c \mu_0} E_{\text{rms}}^2, \quad (33-26)$$

in which $E_{\text{rms}} = E_m / \sqrt{2}$. A *point source* of electromagnetic waves emits the waves *isotropically*—that is, with equal intensity in all directions. The intensity of the waves at distance r from a point source of power P_s is

$$I = \frac{P_s}{4\pi r^2}. \quad (33-27)$$

Radiation Pressure When a surface intercepts electromagnetic radiation, a force and a pressure are exerted on the surface. If the radiation is totally absorbed by the surface, the force is

$$F = \frac{IA}{c} \quad (\text{total absorption}), \quad (33-32)$$

in which I is the intensity of the radiation and A is the area of the surface perpendicular to the path of the radiation. If the radiation is totally reflected back along its original path, the force is

$$F = \frac{2IA}{c} \quad (\text{total reflection back along path}). \quad (33-33)$$

The radiation pressure p_r is the force per unit area:

$$p_r = \frac{I}{c} \quad (\text{total absorption}) \quad (33-34)$$

and $p_r = \frac{2I}{c}$ (total reflection back along path). (33-35)

Polarization Electromagnetic waves are **polarized** if their electric field vectors are all in a single plane, called the *plane of oscillation*. Light waves from common sources are not polarized; that is, they are **unpolarized**, or **polarized randomly**.

Polarizing Sheets When a polarizing sheet is placed in the path of light, only electric field components of the light parallel to the sheet's **polarizing direction** are *transmitted* by the sheet; components perpendicular to the polarizing direction are absorbed. The light that emerges from a polarizing sheet is polarized parallel to the polarizing direction of the sheet.

If the original light is initially unpolarized, the transmitted intensity I is half the original intensity I_0 :

$$I = \frac{1}{2} I_0. \quad (33-36)$$

If the original light is initially polarized, the transmitted intensity depends on the angle θ between the polarization direction of the original light and the polarizing direction of the sheet:

$$I = I_0 \cos^2 \theta. \quad (33-38)$$

Geometrical Optics *Geometrical optics* is an approximate treatment of light in which light waves are represented as straight-line rays.

** View All Solutions Here **

Q U E S T I O N S

1 If the magnetic field of a light wave oscillates parallel to a y axis and is given by $B_y = B_m \sin(kz - \omega t)$, (a) in what direction does the wave travel and (b) parallel to which axis does the associated electric field oscillate?

2 Suppose we rotate the second sheet in Fig. 33-15a, starting with the polarization direction aligned with the y axis ($\theta = 0^\circ$) and ending with it aligned with the x axis ($\theta = 90^\circ$). Which of the four curves in Fig. 33-26 best shows the intensity of the light through the three-sheet system during this 90° rotation?

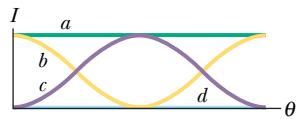


Fig. 33-26 Question 2.

3 (a) Figure 33-27 shows light reaching a polarizing sheet whose polarizing direction is parallel to a y axis. We shall rotate the sheet 40° clockwise about the light's indicated line of travel. During this rotation, does the fraction of the initial light intensity passed by the

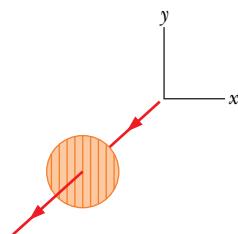


Fig. 33-27 Question 3.

Reflection and Refraction When a light ray encounters a boundary between two transparent media, a **reflected** ray and a **refracted** ray generally appear. Both rays remain in the plane of incidence. The **angle of reflection** is equal to the angle of incidence, and the **angle of refraction** is related to the angle of incidence by Snell's law,

$$n_2 \sin \theta_2 = n_1 \sin \theta_1 \quad (\text{refraction}), \quad (33-40)$$

where n_1 and n_2 are the indexes of refraction of the media in which the incident and refracted rays travel.

Total Internal Reflection A wave encountering a boundary across which the index of refraction decreases will experience **total internal reflection** if the angle of incidence exceeds a **critical angle** θ_c , where

$$\theta_c = \sin^{-1} \frac{n_2}{n_1} \quad (\text{critical angle}). \quad (33-45)$$

Polarization by Reflection A reflected wave will be fully **polarized**, with its \vec{E} vectors perpendicular to the plane of incidence, if it strikes a boundary at the **Brewster angle** θ_B , where

$$\theta_B = \tan^{-1} \frac{n_2}{n_1} \quad (\text{Brewster angle}). \quad (33-49)$$

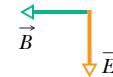


Fig. 33-28
Question 4.

sheet increase, decrease, or remain the same if the light is (a) initially unpolarized, (b) initially polarized parallel to the x axis, and (c) initially polarized parallel to the y axis?

4 Figure 33-28 shows the electric and magnetic fields of an electromagnetic wave at a certain instant. Is the wave traveling into the page or out of it?

5 In Fig. 33-15a, start with light that is initially polarized parallel to the x axis, and write the ratio of its final intensity I_3 to its initial intensity I_0 as $I_3/I_0 = A \cos^n \theta$. What are A , n , and θ if we rotate the polarizing direction of the first sheet (a) 60° counterclockwise and (b) 90° clockwise from what is shown?

6 In Fig. 33-29, unpolarized light is sent into a system of five polarizing sheets. Their polarizing directions, measured counterclockwise from the positive direction of the y axis, are the following: sheet 1, 35° ; sheet 2, 0° ; sheet 3, 0° ; sheet 4, 110° ; sheet 5, 45° . Sheet 3 is then rotated 180° counterclockwise about the light ray. During

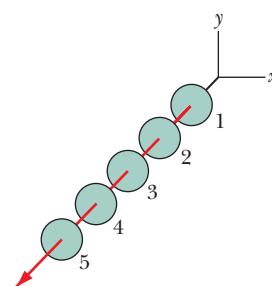


Fig. 33-29 Question 6.

** View All Solutions Here **

that rotation, at what angles (measured counterclockwise from the y axis) is the transmission of light through the system eliminated?

- 7** Figure 33-30 shows rays of monochromatic light passing through three materials a , b , and c . Rank the materials according to the index of refraction, greatest first.

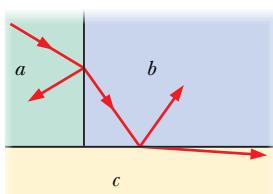


Fig. 33-30 Question 7.

- 8** Figure 33-31 shows the multiple reflections of a light ray along a glass corridor where the walls are either parallel or perpendicular to one another. If the angle of incidence at point a is 30° , what are the angles of reflection of the light ray at points b , c , d , e , and f ?

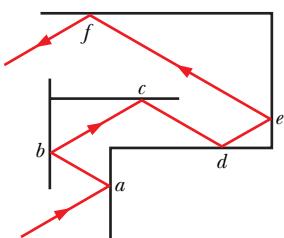


Fig. 33-31 Question 8.

- 9** Figure 33-32 shows four long horizontal layers $A–D$ of different materials, with air above and below them. The index of refraction of each material is given. Rays of light are sent into the left end of each layer as shown. In which layer is there the possibility of totally trapping the light in that layer so that, after many reflections, all the light reaches the right end of the layer?

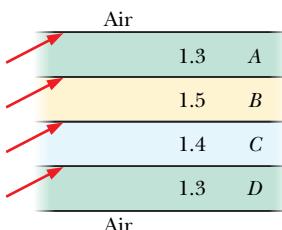


Fig. 33-32 Question 9.

- 10** The leftmost block in Fig. 33-33 depicts total internal reflection for light inside a material with an index of refraction n_1 when air is outside the material. A light ray reaching point A from

anywhere within the shaded region at the left (such as the ray shown) fully reflects at that point and ends up in the shaded region at the right. The other blocks show similar situations for two other materials. Rank the indexes of refraction of the three materials, greatest first.

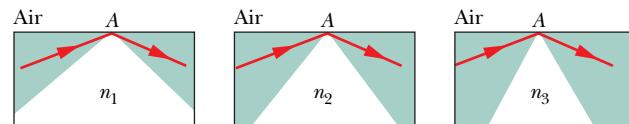


Fig. 33-33 Question 10.

- 11** Each part of Fig. 33-34 shows light that refracts through an interface between two materials. The incident ray (shown gray in the figure) consists of red and blue light. The approximate index of refraction for visible light is indicated for each material. Which of the three parts show physically possible refraction? (Hint: First consider the refraction in general, regardless of the color, and then consider how red and blue light refract differently.)

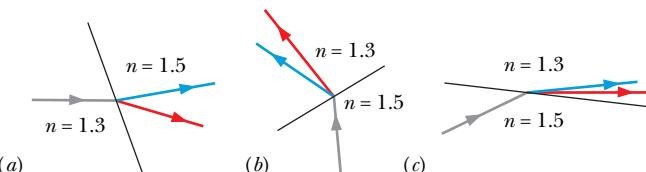


Fig. 33-34 Question 11.

- 12** In Fig. 33-35, light travels from material a , through three layers of other materials with surfaces parallel to one another, and then back into another layer of material a . The refractions (but not the associated reflections) at the surfaces are shown. Rank the materials according to index of refraction, greatest first. (Hint: The parallel arrangement of the surfaces allows comparison.)

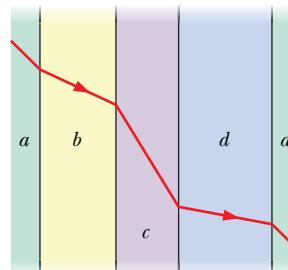


Fig. 33-35 Question 12.



Tutoring problem available (at instructor's discretion) in *WileyPLUS* and *WebAssign*

SSM Worked-out solution available in *Student Solutions Manual*

• • • Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>



sec. 33-2 Maxwell's Rainbow

- 1** A certain helium–neon laser emits red light in a narrow band of wavelengths centered at 632.8 nm and with a “wavelength width” (such as on the scale of Fig. 33-1) of 0.0100 nm. What is the corresponding “frequency width” for the emission?

- 2** Project Seafarer was an ambitious program to construct an enormous antenna, buried underground on a site about $10\ 000\ \text{km}^2$ in area. Its purpose was to transmit signals to submarines while they were deeply submerged. If the effective wavelength

were 1.0×10^4 Earth radii, what would be the (a) frequency and (b) period of the radiations emitted? Ordinarily, electromagnetic radiations do not penetrate very far into conductors such as seawater, and so normal signals cannot reach the submarines.

- 3** From Fig. 33-2, approximate the (a) smaller and (b) larger wavelength at which the eye of a standard observer has half the eye’s maximum sensitivity. What are the (c) wavelength, (d) frequency, and (e) period of the light at which the eye is the most sensitive?

916 CHAPTER 33 ELECTROMAGNETIC WAVES

- 4 About how far apart must you hold your hands for them to be separated by 1.0 nano-light-second (the distance light travels in 1.0 ns)?

sec. 33-3 The Traveling Electromagnetic Wave, Qualitatively

- 5 **SSM** What inductance must be connected to a 17 pF capacitor in an oscillator capable of generating 550 nm (i.e., visible) electromagnetic waves? Comment on your answer.

- 6 What is the wavelength of the electromagnetic wave emitted by the oscillator–antenna system of Fig. 33-3 if $L = 0.253 \mu\text{H}$ and $C = 25.0 \text{ pF}$?

sec. 33-5 Energy Transport and the Poynting Vector

- 7 What is the intensity of a traveling plane electromagnetic wave if B_m is $1.0 \times 10^{-4} \text{ T}$?

- 8 Assume (unrealistically) that a TV station acts as a point source broadcasting isotropically at 1.0 MW. What is the intensity of the transmitted signal reaching Proxima Centauri, the star nearest our solar system, 4.3 ly away? (An alien civilization at that distance might be able to watch *X Files*.) A light-year (ly) is the distance light travels in one year.

- 9 **ILW** Some neodymium–glass lasers can provide 100 TW of power in 1.0 ns pulses at a wavelength of $0.26 \mu\text{m}$. How much energy is contained in a single pulse?

- 10 A plane electromagnetic wave has a maximum electric field magnitude of $3.20 \times 10^{-4} \text{ V/m}$. Find the magnetic field amplitude.

- 11 **ILW** A plane electromagnetic wave traveling in the positive direction of an x axis in vacuum has components $E_x = E_y = 0$ and $E_z = (2.0 \text{ V/m}) \cos[(\pi \times 10^{15} \text{ s}^{-1})(t - x/c)]$. (a) What is the amplitude of the magnetic field component? (b) Parallel to which axis does the magnetic field oscillate? (c) When the electric field component is in the positive direction of the z axis at a certain point P , what is the direction of the magnetic field component there?

- 12 In a plane radio wave the maximum value of the electric field component is 5.00 V/m. Calculate (a) the maximum value of the magnetic field component and (b) the wave intensity.

- 13 Sunlight just outside Earth's atmosphere has an intensity of 1.40 kW/m^2 . Calculate (a) E_m and (b) B_m for sunlight there, assuming it to be a plane wave.

- 14 **GO** An isotropic point source emits light at wavelength 500 nm, at the rate of 200 W. A light detector is positioned 400 m from the source. What is the maximum rate $\partial B/\partial t$ at which the magnetic component of the light changes with time at the detector's location?

- 15 An airplane flying at a distance of 10 km from a radio transmitter receives a signal of intensity $10 \mu\text{W/m}^2$. What is the amplitude of the (a) electric and (b) magnetic component of the signal at the airplane? (c) If the transmitter radiates uniformly over a hemisphere, what is the transmission power?

- 16 Frank D. Drake, an investigator in the SETI (Search for Extra-Terrestrial Intelligence) program, once said that the large radio telescope in Arecibo, Puerto Rico (Fig. 33-36), "can detect a signal which lays down on the entire surface of the earth a power of only one picowatt." (a) What is the power that would be received by the Arecibo antenna for such a signal? The antenna diameter is 300 m. (b) What would be the power of an isotropic

source at the center of our galaxy that could provide such a signal? The galactic center is 2.2×10^4 ly away. A light-year is the distance light travels in one year.



Fig. 33-36 Problem 16. Radio telescope at Arecibo. (Courtesy Cornell University)

- 17 The maximum electric field 10 m from an isotropic point source of light is 2.0 V/m. What are (a) the maximum value of the magnetic field and (b) the average intensity of the light there? (c) What is the power of the source?

- 18 The intensity I of light from an isotropic point source is determined as a function of distance r from the source. Figure 33-37 gives intensity I versus the inverse square r^{-2} of that distance. The vertical axis scale is set by $I_s = 200 \text{ W/m}^2$, and the horizontal axis scale is set by $r_s^{-2} = 8.0 \text{ m}^{-2}$. What is the power of the source?

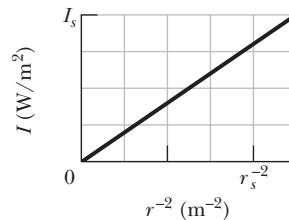


Fig. 33-37 Problem 18.

sec. 33-6 Radiation Pressure

- 19 **SSM** High-power lasers are used to compress a plasma (a gas of charged particles) by radiation pressure. A laser generating radiation pulses with peak power $1.5 \times 10^3 \text{ MW}$ is focused onto 1.0 mm^2 of high-electron-density plasma. Find the pressure exerted on the plasma if the plasma reflects all the light beams directly back along their paths.

- 20 Radiation from the Sun reaching Earth (just outside the atmosphere) has an intensity of 1.4 kW/m^2 . (a) Assuming that Earth (and its atmosphere) behaves like a flat disk perpendicular to the Sun's rays and that all the incident energy is absorbed, calculate the force on Earth due to radiation pressure. (b) For comparison, calculate the force due to the Sun's gravitational attraction.

- 21 **ILW** What is the radiation pressure 1.5 m away from a 500 W lightbulb? Assume that the surface on which the pressure is exerted faces the bulb and is perfectly absorbing and that the bulb radiates uniformly in all directions.

- 22 A black, totally absorbing piece of cardboard of area $A = 2.0 \text{ cm}^2$ intercepts light with an intensity of 10 W/m^2 from a camera strobe light. What radiation pressure is produced on the cardboard by the light?

- 23 Someone plans to float a small, totally absorbing sphere 0.500 m above an isotropic point source of light, so that the upward ra-

PROBLEMS

917

diation force from the light matches the downward gravitational force on the sphere. The sphere's density is 19.0 g/cm^3 , and its radius is 2.00 mm. (a) What power would be required of the light source? (b) Even if such a source were made, why would the support of the sphere be unstable?

••24 GO It has been proposed that a spaceship might be propelled in the solar system by radiation pressure, using a large sail made of foil. How large must the surface area of the sail be if the radiation force is to be equal in magnitude to the Sun's gravitational attraction? Assume that the mass of the ship + sail is 1500 kg, that the sail is perfectly reflecting, and that the sail is oriented perpendicular to the Sun's rays. See Appendix C for needed data. (With a larger sail, the ship is continuously driven away from the Sun.)

••25 SSM Prove, for a plane electromagnetic wave that is normally incident on a flat surface, that the radiation pressure on the surface is equal to the energy density in the incident beam. (This relation between pressure and energy density holds no matter what fraction of the incident energy is reflected.)

••26 In Fig. 33-38, a laser beam of power 4.60 W and diameter $D = 2.60 \text{ mm}$ is directed upward at one circular face (of diameter $d < 2.60 \text{ mm}$) of a perfectly reflecting cylinder. The cylinder is levitated because the upward radiation force matches the downward gravitational force. If the cylinder's density is 1.20 g/cm^3 , what is its height H ?

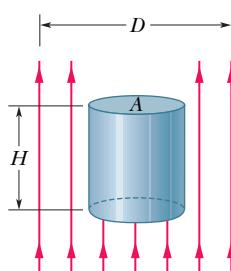


Fig. 33-38
Problem 26.

••27 SSM WWW A plane electromagnetic wave, with wavelength 3.0 m, travels in vacuum in the positive direction of an x axis. The electric field, of amplitude 300 V/m, oscillates parallel to the y axis. What are the (a) frequency, (b) angular frequency, and (c) angular wave number of the wave? (d) What is the amplitude of the magnetic field component? (e) Parallel to which axis does the magnetic field oscillate? (f) What is the time-averaged rate of energy flow in watts per square meter associated with this wave? The wave uniformly illuminates a surface of area 2.0 m^2 . If the surface totally absorbs the wave, what are (g) the rate at which momentum is transferred to the surface and (h) the radiation pressure on the surface?

••28 The average intensity of the solar radiation that strikes normally on a surface just outside Earth's atmosphere is 1.4 kW/m^2 . (a) What radiation pressure p_r is exerted on this surface, assuming complete absorption? (b) For comparison, find the ratio of p_r to Earth's sea-level atmospheric pressure, which is $1.0 \times 10^5 \text{ Pa}$.

••29 SSM A small spaceship with a mass of only $1.5 \times 10^3 \text{ kg}$ (including an astronaut) is drifting in outer space with negligible gravitational forces acting on it. If the astronaut turns on a 10 kW laser beam, what speed will the ship attain in 1.0 day because of the momentum carried away by the beam?

••30 A small laser emits light at power 5.00 mW and wavelength 633 nm. The laser beam is focused (narrowed) until its diameter matches the 1266 nm diameter of a sphere placed in its path. The sphere is perfectly absorbing and has density $5.00 \times 10^3 \text{ kg/m}^3$. What are (a) the beam intensity at the sphere's location, (b) the radiation pressure on the sphere, (c) the magnitude of the corresponding force, and (d) the magnitude of the acceleration that force alone would give the sphere?

••31 As a comet swings around the Sun, ice on the comet's surface vaporizes, releasing trapped dust particles and ions. The ions, because they are electrically charged, are forced by the electrically charged *solar wind* into a straight *ion tail* that points radially away from the Sun (Fig. 33-39). The (electrically neutral) dust particles are pushed radially outward from the Sun by the radiation force on them from sunlight. Assume that the dust particles are spherical, have density $3.5 \times 10^3 \text{ kg/m}^3$, and are totally absorbing. (a) What radius must a particle have in order to follow a straight path, like path 2 in the figure? (b) If its radius is larger, does its path curve away from the Sun (like path 1) or toward the Sun (like path 3)?

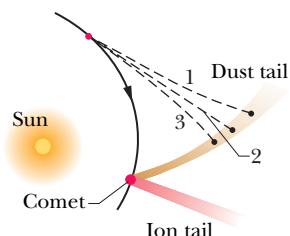


Fig. 33-39 Problem 31.

sec. 33-7 Polarization

••32 In Fig. 33-40, initially unpolarized light is sent into a system of three polarizing sheets whose polarizing directions make angles of $\theta_1 = \theta_2 = \theta_3 = 50^\circ$ with the direction of the y axis. What percentage of the initial intensity is transmitted by the system? (Hint: Be careful with the angles.)

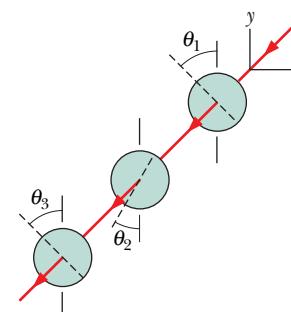


Fig. 33-40 Problems 32 and 33.

••33 SSM In Fig. 33-40, initially unpolarized light is sent into a system of three polarizing sheets whose polarizing directions make angles of $\theta_1 = 40^\circ$, $\theta_2 = 20^\circ$, and $\theta_3 = 40^\circ$ with the direction of the y axis. What percentage of the light's initial intensity is transmitted by the system? (Hint: Be careful with the angles.)

••34 GO In Fig. 33-41, a beam of unpolarized light, with intensity 43 W/m^2 , is sent into a system of two polarizing sheets with polarizing directions at angles $\theta_1 = 70^\circ$ and $\theta_2 = 90^\circ$ to the y axis. What is the intensity of the light transmitted by the system?

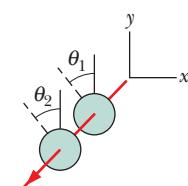


Fig. 33-41
Problems 34, 35,
and 42.

••35 ILW In Fig. 33-41, a beam of light, with intensity 43 W/m^2 and polarization parallel to a y axis, is sent into a system of two polarizing sheets with polarizing directions at angles of $\theta_1 = 70^\circ$ and $\theta_2 = 90^\circ$ to the y axis. What is the intensity of the light transmitted by the two-sheet system?

••36 At a beach the light is generally partially polarized due to reflections off sand and water. At a particular beach on a particular day near sundown, the horizontal component of the electric field vector is 2.3 times the vertical component. A standing sunbather puts on polarizing sunglasses; the glasses eliminate the

918 CHAPTER 33 ELECTROMAGNETIC WAVES

horizontal field component. (a) What fraction of the light intensity received before the glasses were put on now reaches the sunbather's eyes? (b) The sunbather, still wearing the glasses, lies on his side. What fraction of the light intensity received before the glasses were put on now reaches his eyes?

••37 [SSM WWW] We want to rotate the direction of polarization of a beam of polarized light through 90° by sending the beam through one or more polarizing sheets. (a) What is the minimum number of sheets required? (b) What is the minimum number of sheets required if the transmitted intensity is to be more than 60% of the original intensity?

••38 In Fig. 33-42, unpolarized light is sent into a system of three polarizing sheets. The angles θ_1 , θ_2 , and θ_3 of the polarizing directions are measured counterclockwise from the positive direction of the y axis (they are not drawn to scale). Angles θ_1 and θ_3 are fixed, but angle θ_2 can be varied. Figure 33-43 gives the intensity of the light emerging from sheet 3 as a function of θ_2 . (The scale of the intensity axis is not indicated.) What percentage of the light's initial intensity is transmitted by the system when $\theta_2 = 30^\circ$?

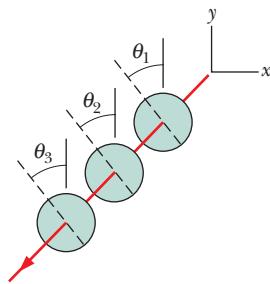


Fig. 33-42 Problems 38, 40, and 44.

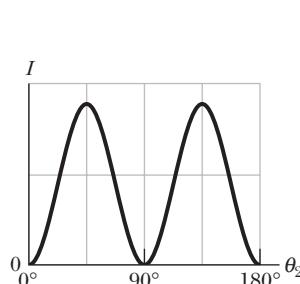


Fig. 33-43 Problem 38.

••39 Unpolarized light of intensity 10 mW/m^2 is sent into a polarizing sheet as in Fig. 33-11. What are (a) the amplitude of the electric field component of the transmitted light and (b) the radiation pressure on the sheet due to its absorbing some of the light?

••40 In Fig. 33-42, unpolarized light is sent into a system of three polarizing sheets. The angles θ_1 , θ_2 , and θ_3 of the polarizing directions are measured counterclockwise from the positive direction of the y axis (they are not drawn to scale). Angles θ_1 and θ_3 are fixed, but angle θ_2 can be varied. Figure 33-44 gives the intensity of the light emerging from sheet 3 as a function of θ_2 . (The scale of the intensity axis is not indicated.) What percentage of the light's initial intensity is transmitted by the three-sheet system when $\theta_2 = 90^\circ$?

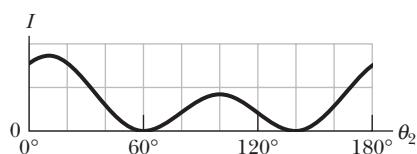


Fig. 33-44 Problem 40.

••41 A beam of polarized light is sent into a system of two polarizing sheets. Relative to the polarization direction of that incident light, the polarizing directions of the sheets are at angles θ for the first sheet and 90° for the second sheet. If 0.10 of the incident intensity is transmitted by the two sheets, what is θ ?

••42 In Fig. 33-41, unpolarized light is sent into a system of two polarizing sheets. The angles θ_1 and θ_2 of the polarizing directions of the sheets are measured counterclockwise from the positive direction of the y axis (they are not drawn to scale in the figure). Angle θ_1 is fixed but angle θ_2 can be varied. Figure 33-45 gives the intensity of the light emerging from sheet 2 as a function of θ_2 . (The scale of the intensity axis is not indicated.) What percentage of the light's initial intensity is transmitted by the two-sheet system when $\theta_2 = 90^\circ$?

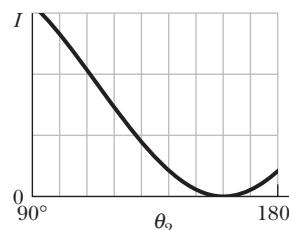


Fig. 33-45 Problem 42.

••43 A beam of partially polarized light can be considered to be a mixture of polarized and unpolarized light. Suppose we send such a beam through a polarizing filter and then rotate the filter through 360° while keeping it perpendicular to the beam. If the transmitted intensity varies by a factor of 5.0 during the rotation, what fraction of the intensity of the original beam is associated with the beam's polarized light?

••44 In Fig. 33-42, unpolarized light is sent into a system of three polarizing sheets, which transmits 0.0500 of the initial light intensity. The polarizing directions of the first and third sheets are at angles $\theta_1 = 0^\circ$ and $\theta_3 = 90^\circ$. What are the (a) smaller and (b) larger possible values of angle θ_2 ($< 90^\circ$) for the polarizing direction of sheet 2?

sec. 33-8 Reflection and Refraction

•45 When the rectangular metal tank in Fig. 33-46 is filled to the top with an unknown liquid, observer O , with eyes level with the top of the tank, can just see corner E . A ray that refracts toward O at the top surface of the liquid is shown. If $D = 85.0 \text{ cm}$ and $L = 1.10 \text{ m}$, what is the index of refraction of the liquid?

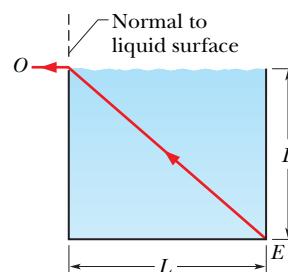


Fig. 33-46 Problem 45.

•46 In Fig. 33-47a, a light ray in an underlying material is incident at angle θ_1 on a boundary with water, and some of the light refracts into the water. There are two choices of underlying material. For each, the angle of refraction θ_2 versus the incident angle θ_1 is given in Fig. 33-47b. The horizontal axis scale is set by $\theta_{1s} = 90^\circ$. Without calculation, determine whether the index of refraction of (a) material 1 and (b) material 2 is greater or less

PROBLEMS

919

than the index of water ($n = 1.33$). What is the index of refraction of (c) material 1 and (d) material 2?

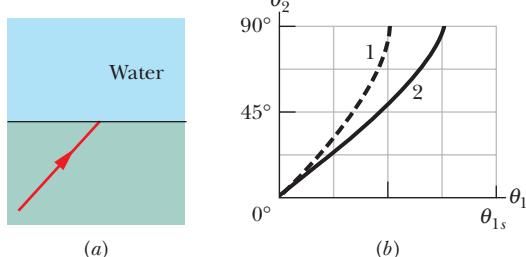


Fig. 33-47 Problem 46.

- 47 Light in vacuum is incident on the surface of a glass slab. In the vacuum the beam makes an angle of 32.0° with the normal to the surface, while in the glass it makes an angle of 21.0° with the normal. What is the index of refraction of the glass?

- 48 In Fig. 33-48a, a light ray in water is incident at angle θ_1 on a boundary with an underlying material, into which some of the light refracts. There are two choices of underlying material. For each, the angle of refraction θ_2 versus the incident angle θ_1 is given in Fig. 33-48b. The vertical axis scale is set by $\theta_{2s} = 90^\circ$. Without calculation, determine whether the index of refraction of (a) material 1 and (b) material 2 is greater or less than the index of water ($n = 1.33$). What is the index of refraction of (c) material 1 and (d) material 2?

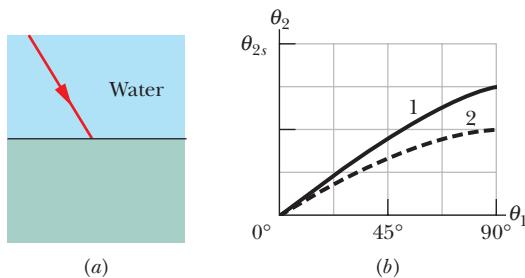


Fig. 33-48 Problem 48.

- 49 Figure 33-49 shows light reflecting from two perpendicular reflecting surfaces *A* and *B*. Find the angle between the incoming ray *i* and the outgoing ray *r'*.

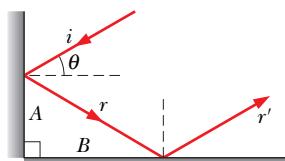


Fig. 33-49 Problem 49.

- 50 In Fig. 33-50a, a beam of light in material 1 is incident on a boundary at an angle $\theta_1 = 40^\circ$. Some of the light travels through material 2, and then some of it emerges into material 3. The two boundaries between the three materials are parallel. The final direction of the beam depends, in part, on the index of refraction n_3 of the third material. Figure 33-50b gives the angle of refraction θ_3 versus n_3 for a range of possible n_3 values. The vertical axis scale is set by $\theta_{3a} = 30.0^\circ$ and $\theta_{3b} = 50.0^\circ$. (a) What is the index

of refraction of material 1, or is the index impossible to calculate without more information? (b) What is the index of refraction of material 2, or is the index impossible to calculate without more information? (c) If θ_1 is changed to 70° and the index of refraction of material 3 is 2.4, what is θ_3 ?

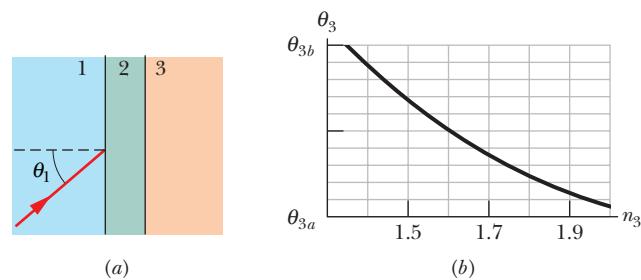


Fig. 33-50 Problem 50.

- 51 In Fig. 33-51, light is incident at angle $\theta_1 = 40.1^\circ$ on a boundary between two transparent materials. Some of the light travels down through the next three layers of transparent materials, while some of it reflects upward and then escapes into the air. If $n_1 = 1.30$, $n_2 = 1.40$, $n_3 = 1.32$, and $n_4 = 1.45$, what is the value of (a) θ_5 in the air and (b) θ_4 in the bottom material?

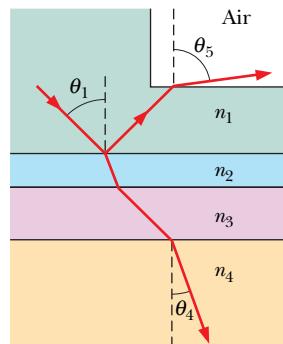


Fig. 33-51 Problem 51.

- 52 In Fig. 33-52a, a beam of light in material 1 is incident on a boundary at an angle of $\theta_1 = 30^\circ$. The extent of refraction of the light into material 2 depends, in part, on the index of refraction n_2 of material 2. Figure 33-52b gives the angle of refraction θ_2 versus n_2 for a range of possible n_2 values. The vertical axis scale is set by $\theta_{2a} = 20.0^\circ$ and $\theta_{2b} = 40.0^\circ$. (a) What is the index of refraction of material 1? (b) If the incident angle is changed to 60° and material 2 has $n_2 = 2.4$, then what is angle θ_2 ?

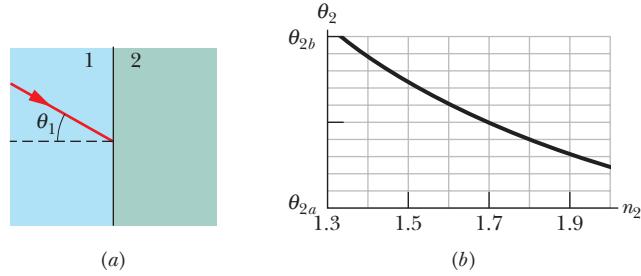


Fig. 33-52 Problem 52.

••53 SSM ILW WWW In Fig. 33-53, a ray is incident on one face of a triangular glass prism in air. The angle of incidence θ is chosen so that the emerging ray also makes the same angle θ with the normal to the other face. Show that the index of refraction n of the glass prism is given by

$$n = \frac{\sin \frac{1}{2}(\psi + \phi)}{\sin \frac{1}{2}\phi},$$

where ϕ is the vertex angle of the prism and ψ is the *deviation angle*, the total angle through which the beam is turned in passing through the prism. (Under these conditions the deviation angle ψ has the smallest possible value, which is called the *angle of minimum deviation*.)

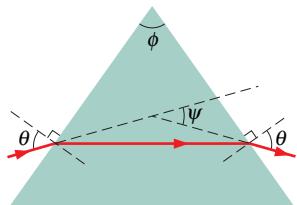


Fig. 33-53
Problems 53 and 64.

••54 Dispersion in a window pane. In Fig. 33-54, a beam of white light is incident at angle $\theta = 50^\circ$ on a common window pane (shown in cross section). For the pane's type of glass, the index of refraction for visible light ranges from 1.524 at the blue end of the spectrum to 1.509 at the red end. The two sides of the pane are parallel. What is the angular spread of the colors in the beam (a) when the light enters the pane and (b) when it emerges from the opposite side? (Hint: When you look at an object through a window pane, are the colors in the light from the object dispersed as shown in, say, Fig. 33-20?)

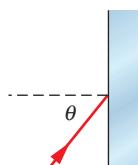


Fig. 33-54
Problem 54.

••55 GO SSM In Fig. 33-55, a 2.00-m-long vertical pole extends from the bottom of a swimming pool to a point 50.0 cm above the water. Sunlight is incident at angle $\theta = 55.0^\circ$. What is the length of the shadow of the pole on the level bottom of the pool?

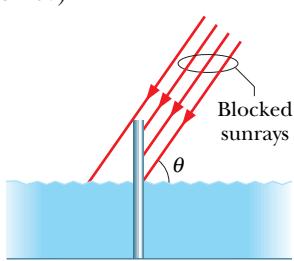


Fig. 33-55
Problem 55.

••56 Rainbows from square drops. Suppose that, on some surreal world, raindrops had a square cross section and always fell with one face horizontal. Figure 33-56 shows such a falling drop, with a white beam of sunlight incident at $\theta = 70.0^\circ$ at point P. The part of the light that enters the drop then travels to point A, where some of it refracts out into the air and the rest reflects. That reflected light then travels to point B, where again some of the light refracts out into the air and the rest reflects. What is the difference in the angles of the red light ($n = 1.331$) and the blue light ($n = 1.343$) that emerge at

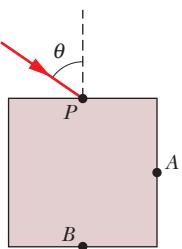


Fig. 33-56
Problem 56.

(a) point A and (b) point B? (This angular difference in the light emerging at, say, point A would be the rainbow's angular width.)

sec. 33-9 Total Internal Reflection

••57 A point source of light is 80.0 cm below the surface of a body of water. Find the diameter of the circle at the surface through which light emerges from the water.

••58 The index of refraction of benzene is 1.8. What is the critical angle for a light ray traveling in benzene toward a flat layer of air above the benzene?

••59 SSM ILW In Fig. 33-57, a ray of light is perpendicular to the face ab of a glass prism ($n = 1.52$). Find the largest value for the angle ϕ so that the ray is totally reflected at face ac if the prism is immersed (a) in air and (b) in water.



Fig. 33-57 Problem 59.

••60 In Fig. 33-58, light from ray A refracts from material 1 ($n_1 = 1.60$) into a thin layer of material 2 ($n_2 = 1.80$), crosses that layer, and is then incident at the critical angle on the interface between materials 2 and 3 ($n_3 = 1.30$). (a) What is the value of incident angle θ_A ? (b) If θ_A is decreased, does part of the light refract into material 3?

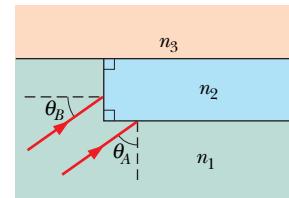


Fig. 33-58 Problem 60.

Light from ray B refracts from material 1 into the thin layer, crosses that layer, and is then incident at the critical angle on the interface between materials 2 and 3. (c) What is the value of incident angle θ_B ? (d) If θ_B is decreased, does part of the light refract into material 3?

••61 GO In Fig. 33-59, light initially in material 1 refracts into material 2, crosses that material, and is then incident at the critical angle on the interface between materials 2 and 3. The indexes of refraction are $n_1 = 1.60$, $n_2 = 1.40$, and $n_3 = 1.20$. (a) What is angle θ ? (b) If θ is increased, is there refraction of light into material 3?

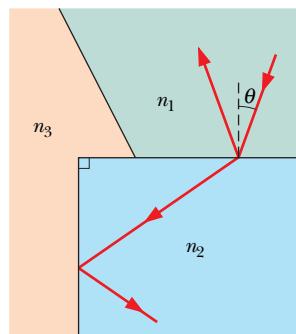


Fig. 33-59 Problem 61.

••62 GO A catfish is 2.00 m below the surface of a smooth lake. (a) What is the diameter of the circle on the surface through which the fish can see the world outside the water? (b) If the fish descends, does the diameter of the circle increase, decrease, or remain the same?

••63 In Fig. 33-60, light enters a 90° triangular prism at point P with incident angle θ , and then some of it refracts at point Q with an angle of refraction of 90°. (a) What is the index of refraction of the prism in terms of θ ? (b) What, numerically, is the maximum value that the index of refraction can have? Does light emerge at Q if the incident angle at P is (c) increased slightly and (d) decreased slightly?

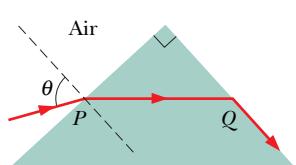


Fig. 33-60 Problem 63.

••64 Suppose the prism of Fig. 33-53 has apex angle $\phi = 60.0^\circ$ and index of refraction $n = 1.60$. (a) What is the smallest angle of incidence θ for which a ray can enter the left face of the prism and exit the right face? (b) What angle of incidence θ is required for the ray to exit the prism with an identical angle θ for its refraction, as it does in Fig. 33-53?

••65 GO Figure 33-61 depicts a simplistic optical fiber: a plastic core ($n_1 = 1.58$) is surrounded by a plastic sheath ($n_2 = 1.53$). A light ray is incident on one end of the fiber at angle θ . The ray is to undergo total internal reflection at point A, where it encounters the core–sheath boundary. (Thus there is no loss of light through that boundary.) What is the maximum value of θ that allows total internal reflection at A?

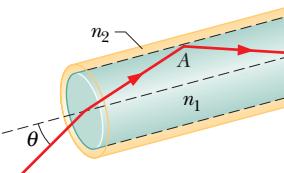


Fig. 33-61 Problem 65.

••66 In Fig. 33-62, a light ray in air is incident at angle θ_1 on a block of transparent plastic with an index of refraction of 1.56. The dimensions indicated are $H = 2.00 \text{ cm}$ and $W = 3.00 \text{ cm}$. The light passes through the block to one of its sides and there undergoes reflection (inside the block) and possibly refraction (out into the air). This is the point of *first reflection*. The reflected light then passes through the block to another of its sides—a point of *second reflection*. If $\theta_1 = 40^\circ$, on which side is the point of (a) first reflection and (b) second reflection? If there is refraction at the point of (c) first reflection and (d) second reflection, give the angle of refraction; if not, answer “none.” If $\theta_1 = 70^\circ$, on which side is the point of (e) first reflection and (f) second reflection? If there is refraction at the point of (g) first reflection and (h) second reflection, give the angle of refraction; if not, answer “none.”

••67 In the ray diagram of Fig. 33-63, where the angles are not drawn to scale, the ray is incident at the critical angle on the interface between materials 2 and 3. Angle $\phi = 60.0^\circ$, and two of the indexes of refraction are $n_1 = 1.70$ and $n_2 = 1.60$. Find (a) index of refraction n_3 and (b) angle θ . (c) If θ is decreased, does light refract into material 3?

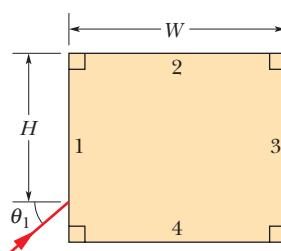


Fig. 33-62 Problem 66.

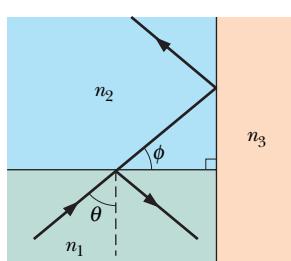


Fig. 33-63 Problem 67.

sec. 33-10 Polarization by Reflection

••68 (a) At what angle of incidence will the light reflected from water be completely polarized? (b) Does this angle depend on the wavelength of the light?

••69 SSM Light that is traveling in water (with an index of refraction of 1.33) is incident on a plate of glass (with index of refraction 1.53). At what angle of incidence does the reflected light end up fully polarized?

••70 In Fig. 33-64, a light ray in air is incident on a flat layer of material 2 that has an index of refraction $n_2 = 1.5$. Beneath material 2 is material 3 with an index of refraction n_3 . The ray is incident on the air–material 2 interface at the Brewster angle for that interface. The ray of light refracted into material 2 happens to be incident on the material 2–material 3 interface at the Brewster angle for that interface. What is the value of n_3 ?

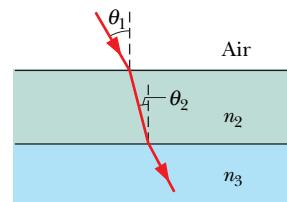


Fig. 33-64 Problem 70.

Additional Problems

71 SSM (a) How long does it take a radio signal to travel 150 km from a transmitter to a receiving antenna? (b) We see a full Moon by reflected sunlight. How much earlier did the light that enters our eye leave the Sun? The Earth–Moon and Earth–Sun distances are $3.8 \times 10^5 \text{ km}$ and $1.5 \times 10^8 \text{ km}$, respectively. (c) What is the round-trip travel time for light between Earth and a spaceship orbiting Saturn, $1.3 \times 10^9 \text{ km}$ distant? (d) The Crab nebula, which is about 6500 light-years (ly) distant, is thought to be the result of a supernova explosion recorded by Chinese astronomers in A.D. 1054. In approximately what year did the explosion actually occur? (When we look into the night sky, we are effectively looking back in time.)

72 An electromagnetic wave with frequency $4.00 \times 10^{14} \text{ Hz}$ travels through vacuum in the positive direction of an x axis. The wave has its electric field directed parallel to the y axis, with amplitude E_m . At time $t = 0$, the electric field at point P on the x axis has a value of $+E_m/4$ and is decreasing with time. What is the distance along the x axis from point P to the first point with $E = 0$ if we search in (a) the negative direction and (b) the positive direction of the x axis?

73 SSM The electric component of a beam of polarized light is

$$E_y = (5.00 \text{ V/m}) \sin[(1.00 \times 10^6 \text{ m}^{-1})z + \omega t].$$

(a) Write an expression for the magnetic field component of the wave, including a value for ω . What are the (b) wavelength, (c) period, and (d) intensity of this light? (e) Parallel to which axis does the magnetic field oscillate? (f) In which region of the electromagnetic spectrum is this wave?

74 A particle in the solar system is under the combined influence of the Sun’s gravitational attraction and the radiation force due to the Sun’s rays. Assume that the particle is a sphere of density $1.0 \times 10^3 \text{ kg/m}^3$ and that all the incident light is absorbed. (a) Show that, if its radius is less than some critical radius R , the particle will be blown out of the solar system. (b) Calculate the critical radius.

75 SSM In Fig. 33-65, a light ray enters a glass slab at point *A* at incident angle $\theta_i = 45.0^\circ$ and then undergoes total internal reflection at point *B*. What minimum value for the index of refraction of the glass can be inferred from this information?

76 GO In Fig. 33-66, unpolarized light with an intensity of 25 W/m^2 is sent into a system of four polarizing sheets with polarizing directions at angles $\theta_1 = 40^\circ$, $\theta_2 = 20^\circ$, $\theta_3 = 20^\circ$, and $\theta_4 = 30^\circ$. What is the intensity of the light that emerges from the system?

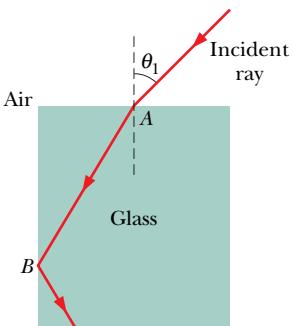


Fig. 33-65 Problem 75.

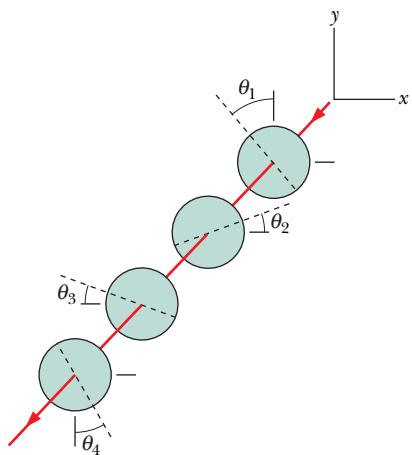


Fig. 33-66 Problem 76.

77 *Rainbow.* Figure 33-67 shows a light ray entering and leaving a falling, spherical raindrop after one internal reflection (see Fig. 33-21a). The final direction of travel is deviated (turned) from the initial direction of travel by angular deviation θ_{dev} . (a) Show that θ_{dev} is

$$\theta_{\text{dev}} = 180^\circ + 2\theta_i - 4\theta_r,$$

where θ_i is the angle of incidence of the ray on the drop and θ_r is the angle of refraction of the ray within the drop. (b) Using Snell's law, substitute for θ_r in terms of θ_i and the index of refraction n of the water. Then, on a graphing calculator or with a computer graphing package, graph θ_{dev} versus θ_i for the range of possible θ_i values and for $n = 1.331$ for red light and $n = 1.333$ for blue light.

The red-light curve and the blue-light curve have different minima, which means that there is a different *angle of minimum deviation* for each color. The light of any given color that leaves the drop at that color's angle of minimum deviation is especially bright because rays bunch up at that angle. Thus, the bright red light leaves the drop at one angle and the bright blue light leaves it at another angle.

Determine the angle of minimum deviation from the θ_{dev} curve for (c) red light and (d) blue light. (e) If these colors form the inner

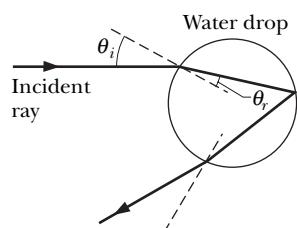


Fig. 33-67 Problem 77.

and outer edges of a rainbow (Fig. 33-21a), what is the angular width of the rainbow?

78 The *primary rainbow* described in Problem 77 is the type commonly seen in regions where rainbows appear. It is produced by light reflecting once inside the drops. Rarer is the *secondary rainbow* described in Section 33-8, produced by light reflecting twice inside the drops (Fig. 33-68a). (a) Show that the angular deviation of light entering and then leaving a spherical water drop is

$$\theta_{\text{dev}} = (180^\circ)k + 2\theta_i - 2(k+1)\theta_r,$$

where k is the number of internal reflections. Using the procedure of Problem 77, find the angle of minimum deviation for (b) red light and (c) blue light in a secondary rainbow. (d) What is the angular width of that rainbow (Fig. 33-21d)?

The *tertiary rainbow* depends on three internal reflections (Fig. 33-68b). It probably occurs but, as noted in Section 33-8, cannot be seen because it is very faint and lies in the bright sky surrounding the Sun. What is the angle of minimum deviation for (e) the red light and (f) the blue light in this rainbow? (g) What is the rainbow's angular width?

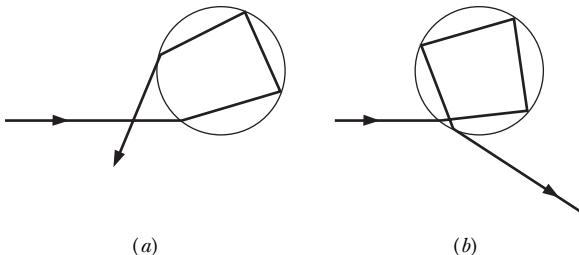


Fig. 33-68 Problem 78.

79 SSM (a) Prove that a ray of light incident on the surface of a sheet of plate glass of thickness t emerges from the opposite face parallel to its initial direction but displaced sideways, as in Fig. 33-69. (b) Show that, for small angles of incidence θ , this displacement is given by

$$x = t\theta \frac{n-1}{n},$$

where n is the index of refraction of the glass and θ is measured in radians.

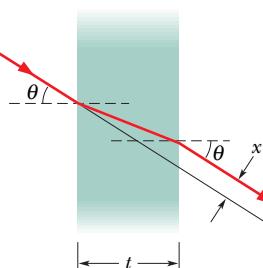


Fig. 33-69 Problem 79.

80 An electromagnetic wave is traveling in the negative y axis. At a particular position and time, the electric field is directed along the positive direction of the z axis and has a magnitude of 100 V/m . What are the (a) magnitude and (b) direction of the corresponding magnetic field?

- 81** The magnetic component of a polarized wave of light is

$$B_x = (4.0 \times 10^{-6} \text{ T}) \sin[(1.57 \times 10^7 \text{ m}^{-1})y + \omega t].$$

(a) Parallel to which axis is the light polarized? What are the (b) frequency and (c) intensity of the light?

- 82** In Fig. 33-70, unpolarized light is sent into the system of three polarizing sheets, where the polarizing directions of the first and third sheets are at angles $\theta_1 = 30^\circ$ (counterclockwise) and $\theta_3 = 30^\circ$ (clockwise). What fraction of the initial light intensity emerges from the system?

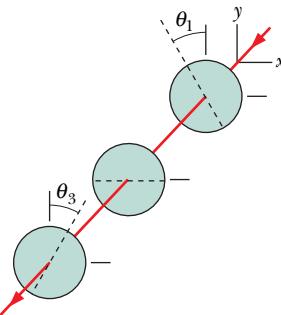


Fig. 33-70 Problem 82.

- 83 SSM** A ray of white light traveling through fused quartz is incident at a quartz–air interface at angle θ_1 . Assume that the index of refraction of quartz is $n = 1.456$ at the red end of the visible range and $n = 1.470$ at the blue end. If θ_1 is (a) 42.00° , (b) 43.10° , and (c) 44.00° , is the refracted light white, white dominated by the red end of the visible range, or white dominated by the blue end of the visible range, or is there no refracted light?

- 84** Three polarizing sheets are stacked. The first and third are crossed; the one between has its polarizing direction at 45.0° to the polarizing directions of the other two. What fraction of the intensity of an originally unpolarized beam is transmitted by the stack?

- 85** In a region of space where gravitational forces can be neglected, a sphere is accelerated by a uniform light beam of intensity 6.0 mW/m^2 . The sphere is totally absorbing and has a radius of $2.0 \mu\text{m}$ and a uniform density of $5.0 \times 10^3 \text{ kg/m}^3$. What is the magnitude of the sphere's acceleration due to the light?

- 86** An unpolarized beam of light is sent into a stack of four polarizing sheets, oriented so that the angle between the polarizing directions of adjacent sheets is 30° . What fraction of the incident intensity is transmitted by the system?

- 87 SSM** During a test, a NATO surveillance radar system, operating at 12 GHz at 180 kW of power, attempts to detect an incoming stealth aircraft at 90 km. Assume that the radar beam is emitted uniformly over a hemisphere. (a) What is the intensity of the beam when the beam reaches the aircraft's location? The aircraft reflects radar waves as though it has a cross-sectional area of only 0.22 m^2 . (b) What is the power of the aircraft's reflection? Assume that the beam is reflected uniformly over a hemisphere. Back at the radar site, what are (c) the intensity, (d) the maximum value of the electric field vector, and (e) the rms value of the magnetic field of the reflected radar beam?

- 88** The magnetic component of an electromagnetic wave in vacuum has an amplitude of 85.8 nT and an angular wave number of 4.00 m^{-1} . What are (a) the frequency of the wave, (b) the rms value of the electric component, and (c) the intensity of the light?

- 89** Calculate the (a) upper and (b) lower limit of the Brewster angle for white light incident on fused quartz. Assume that the wavelength limits of the light are 400 and 700 nm.

- 90** In Fig. 33-71, two light rays pass from air through five layers of transparent plastic and then back into air. The layers have parallel interfaces and unknown thicknesses; their indexes of refraction are $n_1 = 1.7$, $n_2 = 1.6$, $n_3 = 1.5$, $n_4 = 1.4$, and $n_5 = 1.6$. Ray *b* is incident at angle $\theta_b = 20^\circ$. Relative to a normal at the last interface, at what angle do (a) ray *a* and (b) ray *b* emerge? (Hint: Solving the problem algebraically can save time.) If the air at the left and right sides in the figure were, instead, glass with index of refraction 1.5, at what angle would (c) ray *a* and (d) ray *b* emerge?

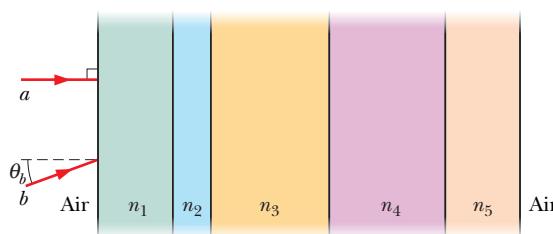


Fig. 33-71 Problem 90.

- 91** A helium-neon laser, radiating at 632.8 nm , has a power output of 3.0 mW . The beam diverges (spreads) at angle $\theta = 0.17 \text{ mrad}$ (Fig. 33-72). (a) What is the intensity of the beam 40 m from the laser? (b) What is the power of a point source providing that intensity at that distance?

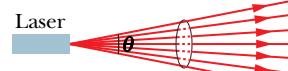


Fig. 33-72 Problem 91.

- 92** In about A.D. 150, Claudius Ptolemy gave the following measured values for the angle of incidence θ_1 and the angle of refraction θ_2 for a light beam passing from air to water:

θ_1	θ_2	θ_1	θ_2
10°	8°	50°	35°
20°	$15^\circ 30'$	60°	$40^\circ 30'$
30°	$22^\circ 30'$	70°	$45^\circ 30'$
40°	29°	80°	50°

Assuming these data are consistent with the law of refraction, use them to find the index of refraction of water. These data are interesting as perhaps the oldest recorded physical measurements.

- 93** A beam of initially unpolarized light is sent through two polarizing sheets placed one on top of the other. What must be the angle between the polarizing directions of the sheets if the intensity of the transmitted light is to be one-third the incident intensity?

34

IMAGES

34-1 WHAT IS PHYSICS?

One goal of physics is to discover the basic laws governing light, such as the law of refraction. A broader goal is to put those laws to use, and perhaps the most important use is the production of images. The first photographic images, made in 1824, were only novelties, but our world now thrives on images. Huge industries are based on the production of images on television, computer, and theater screens. Images from satellites guide military strategists during times of conflict and environmental strategists during times of blight. Camera surveillance can make a subway system more secure, but it can also invade the privacy of unsuspecting citizens. Physiologists and medical engineers are still puzzled by how images are produced by the human eye and the visual cortex of the brain, but they have managed to create mental images in some sightless people by electrical stimulation of the brain's visual cortex.

Our first step in this chapter is to define and classify images. Then we examine several basic ways in which they can be produced.

34-2 Two Types of Image

For you to see, say, a penguin, your eye must intercept some of the light rays spreading from the penguin and then redirect them onto the retina at the rear of the eye. Your visual system, starting with the retina and ending with the visual cortex at the rear of your brain, automatically and subconsciously processes the information provided by the light. That system identifies edges, orientations, textures, shapes, and colors and then rapidly brings to your consciousness an **image** (a reproduction derived from light) of the penguin; you perceive and recognize the penguin as being in the direction from which the light rays came and at the proper distance.

Your visual system goes through this processing and recognition even if the light rays do not come directly from the penguin, but instead reflect toward you from a mirror or refract through the lenses in a pair of binoculars. However, you now see the penguin in the direction from which the light rays came after they reflected or refracted, and the distance you perceive may be quite different from the penguin's true distance.

For example, if the light rays have been reflected toward you from a standard flat mirror, the penguin appears to be behind the mirror because the rays you intercept come from that direction. Of course, the penguin is not back there. This type of image, which is called a **virtual image**, truly exists only within the brain but nevertheless is *said* to exist at the perceived location.

A **real image** differs in that it can be formed on a surface, such as a card or a movie screen. You can see a real image (otherwise movie theaters would be

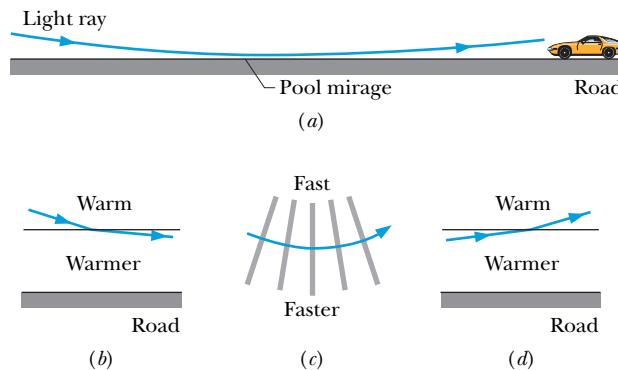


Fig. 34-1 (a) A ray from a low section of the sky refracts through air that is heated by a road (without reaching the road). An observer who intercepts the light perceives it to be from a pool of water on the road. (b) Bending (exaggerated) of a light ray descending across an imaginary boundary from warm air to warmer air. (c) Shifting of wavefronts and associated bending of a ray, which occur because the lower ends of wavefronts move faster in warmer air. (d) Bending of a ray ascending across an imaginary boundary to warm air from warmer air.

empty), but the existence of the image does not depend on your seeing it and it is present even if you are not.

In this chapter we explore several ways in which virtual and real images are formed by reflection (as with mirrors) and refraction (as with lenses). We also distinguish between the two types of image more clearly, but here first is an example of a natural virtual image.

A Common Mirage

A common example of a virtual image is a pool of water that appears to lie on the road some distance ahead of you on a sunny day, but that you can never reach. The pool is a *mirage* (a type of illusion), formed by light rays coming from the low section of the sky in front of you (Fig. 34-1a). As the rays approach the road, they travel through progressively warmer air that has been heated by the road, which is usually relatively warm. With an increase in air temperature, the density of the air—and hence the index of refraction of the air—decreases slightly. Thus, as the rays descend, encountering progressively smaller indexes of refraction, they continuously bend toward the horizontal (Fig. 34-1b).

Once a ray is horizontal, somewhat above the road's surface, it still bends because the lower portion of each associated wavefront is in slightly warmer air and is moving slightly faster than the upper portion of the wavefront (Fig. 34-1c). This nonuniform motion of the wavefronts bends the ray upward. As the ray then ascends, it continues to bend upward through progressively greater indexes of refraction (Fig. 34-1d).

If you intercept some of this light, your visual system automatically infers that it originated along a backward extension of the rays you have intercepted and, to make sense of the light, assumes that it came from the road surface. If the light happens to be bluish from blue sky, the mirage appears bluish, like water. Because the air is probably turbulent due to the heating, the mirage shimmies, as if water waves were present. The bluish coloring and the shimmy enhance the illusion of a pool of water, but you are actually seeing a virtual image of a low section of the sky. As you travel toward the illusionary pool, you no longer intercept the shallow refracted rays and the illusion disappears.

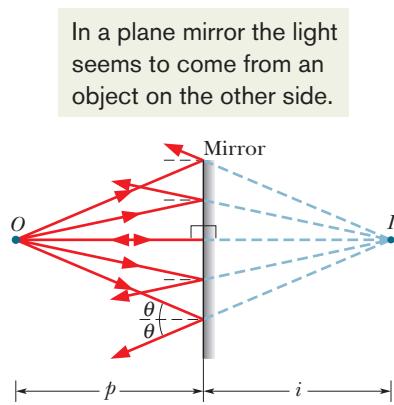


Fig. 34-2 A point source of light O , called the *object*, is a perpendicular distance p in front of a plane mirror. Light rays reaching the mirror from O reflect from the mirror. If your eye intercepts some of the reflected rays, you perceive a point source of light I to be behind the mirror, at a perpendicular distance i . The perceived source I is a virtual image of object O .

34-3 Plane Mirrors

A **mirror** is a surface that can reflect a beam of light in one direction instead of either scattering it widely in many directions or absorbing it. A shiny metal surface acts as a mirror; a concrete wall does not. In this section we examine the images that a **plane mirror** (a flat reflecting surface) can produce.

Figure 34-2 shows a point source of light O , which we shall call the *object*, at a perpendicular distance p in front of a plane mirror. The light that is incident on the mirror is represented with rays spreading from O . The reflection of that light is represented with reflected rays spreading from the mirror. If we extend the reflected rays backward (behind the mirror), we find that the extensions intersect at a point that is a perpendicular distance i behind the mirror.

If you look into the mirror of Fig. 34-2, your eyes intercept some of the reflected light. To make sense of what you see, you perceive a point source of light located at the point of intersection of the extensions. This point source is the image I of object O . It is called a *point image* because it is a point, and it is a virtual image because the rays do not actually pass through it. (As you will see, rays *do* pass through a point of intersection for a real image.)

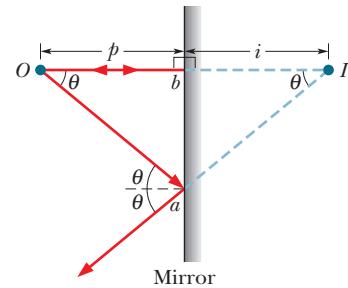


Fig. 34-3 Two rays from Fig. 34-2. Ray Oa makes an arbitrary angle θ with the normal to the mirror surface. Ray Ob is perpendicular to the mirror.

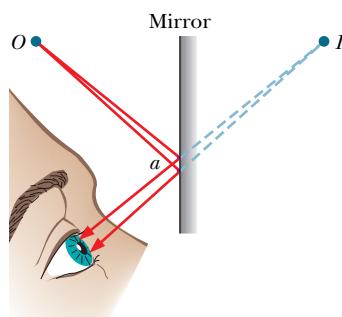


Fig. 34-4 A “pencil” of rays from O enters the eye after reflection at the mirror. Only a small portion of the mirror near point a is involved in this reflection. The light appears to originate at point I behind the mirror.

Figure 34-3 shows two rays selected from the many rays in Fig. 34-2. One reaches the mirror at point b , perpendicularly. The other reaches it at an arbitrary point a , with an angle of incidence θ . The extensions of the two reflected rays are also shown. The right triangles $aOOb$ and aIb have a common side and three equal angles and are thus congruent (equal in size); so their horizontal sides have the same length. That is,

$$Ib = Ob, \quad (34-1)$$

where Ib and Ob are the distances from the mirror to the image and the object, respectively. Equation 34-1 tells us that the image is as far behind the mirror as the object is in front of it. By convention (that is, to get our equations to work out), *object distances* p are taken to be positive quantities and *image distances* i for virtual images (as here) are taken to be negative quantities. Thus, Eq. 34-1 can be written as $|i| = p$ or as

$$i = -p \quad (\text{plane mirror}). \quad (34-2)$$

Only rays that are fairly close together can enter the eye after reflection at a mirror. For the eye position shown in Fig. 34-4, only a small portion of the mirror near point a (a portion smaller than the pupil of the eye) is useful in forming the image. To find this portion, close one eye and look at the mirror image of a small object such as the tip of a pencil. Then move your fingertip over the mirror surface until you cannot see the image. Only that small portion of the mirror under your fingertip produced the image.

Extended Objects

In Fig. 34-5, an extended object O , represented by an upright arrow, is at perpendicular distance p in front of a plane mirror. Each small portion of the object that faces the mirror acts like the point source O of Figs. 34-2 and 34-3. If you intercept the light reflected by the mirror, you perceive a virtual image I that is a composite of the virtual point images of all those portions of the object. This virtual image seems to be at (negative) distance i behind the mirror, with i and p related by Eq. 34-2.

We can also locate the image of an extended object as we did for a point object in Fig. 34-2: we draw some of the rays that reach the mirror from the top of the object, draw the corresponding reflected rays, and then extend those reflected rays behind the mirror until they intersect to form an image of the top of the object. We then do the same for rays from the bottom of the object. As shown in Fig. 34-5, we find that virtual image I has the same orientation and *height* (measured parallel to the mirror) as object O .



Fig. 34-6 A maze of mirrors. (*Courtesy Adrian Fisher. www.mazemaker.com*)

In a plane mirror the image is just as far from the mirror as the object.

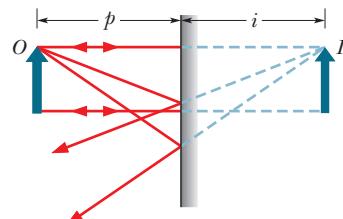
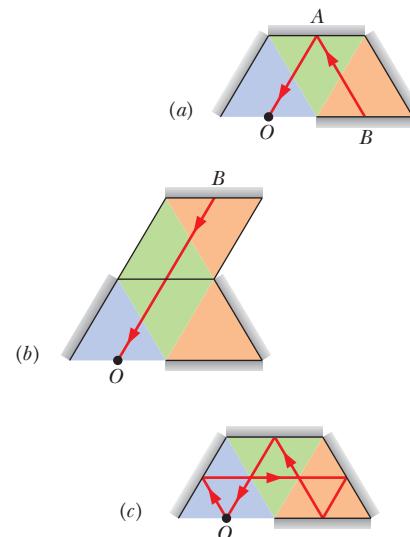


Fig. 34-5 An extended object O and its virtual image I in a plane mirror.



Mirror Maze

In a mirror maze (Fig. 34-6), each wall is covered, floor to ceiling, with a mirror. Walk through such a maze and what you see in most directions is a confusing montage of reflections. In some directions, however, you see a hallway that seems to offer a path through the maze. Take these hallways, though, and you soon learn, after smacking into mirror after mirror, that the hallways are largely an illusion.

Figure 34-7a is an overhead view of a simple mirror maze in which differently painted floor sections form equilateral triangles (60° angles) and walls are covered with vertical mirrors. You look into the maze while standing at point O at the middle of the maze entrance. In most directions, you see a confusing jumble of images. However, you see something curious in the direction of the ray shown in Fig. 34-7a. That ray leaves the middle of mirror B and reflects to you at the middle of mirror A . (The reflection obeys the law of reflection, with the angle of incidence and the angle of reflection both equal to 30° .)

To make sense of the origin of the ray reaching you, your brain automatically extends the ray backward. It appears to originate at a point lying *behind* mirror A . That is, you perceive a virtual image of B behind A , at a distance equal to the actual distance between A and B (Fig. 34-7b). Thus, when you face into the maze in this direction, you see B along an apparent straight hallway consisting of four triangular floor sections.

A hallway seems to lie in front of you.

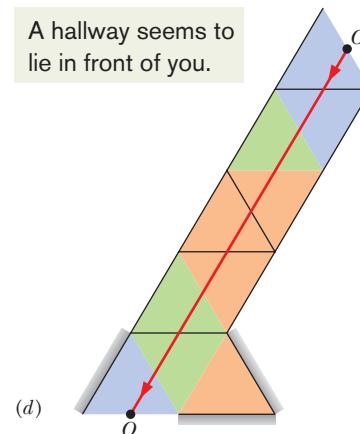
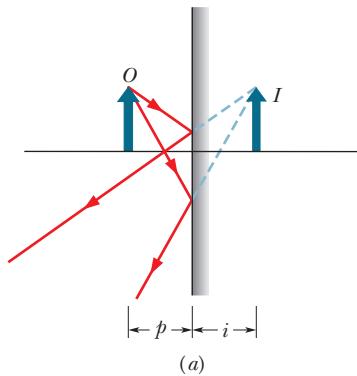
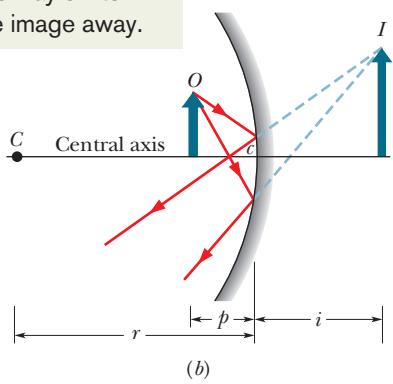


Fig. 34-7 (a) Overhead view of a mirror maze. A ray from mirror B reaches you at O by reflecting from mirror A . (b) Mirror B appears to be behind A . (c) The ray reaching you comes from you. (d) You see a virtual image of yourself at the end of an apparent hallway.

This story is incomplete, however, because the ray reaching you does not originate at mirror *B*—it only reflects there. To find the origin, we continue to apply the law of reflection as we work backwards, reflection by reflection on the mirrors. Working through the four reflections shown in Fig. 34-7c, we finally come to the origin of the ray: you! What you see when you look along the apparent hallway is a virtual image of yourself, at a distance of nine triangular floor sections from you (Fig. 34-7d). (There is a second apparent hallway extending away from point *O*. Which way must you face to look along it?)

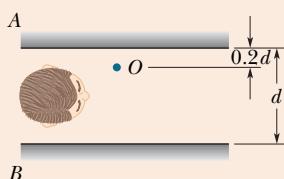


Bending the mirror this way shifts the image away.



CHECKPOINT 1

In the figure you are in a system of two vertical parallel mirrors *A* and *B* separated by distance *d*. A grinning gargoyle is perched at point *O*, a distance $0.2d$ from mirror *A*. Each mirror produces a *first* (least deep) image of the gargoyle. Then each mirror produces a *second* image with the object being the first image in the opposite mirror. Then each mirror produces a *third* image with the object being the second image in the opposite mirror, and so on—you might see hundreds of grinning gargoyle images. How deep behind mirror *A* are the first, second, and third images in mirror *A*?

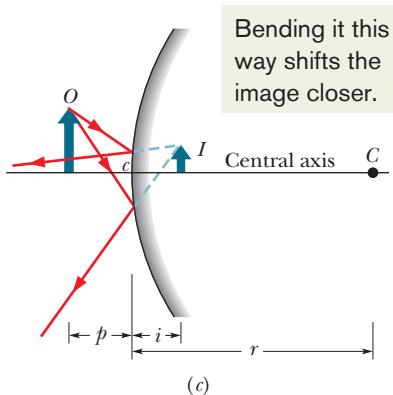


34-4 Spherical Mirrors

We turn now from images produced by plane mirrors to images produced by mirrors with curved surfaces. In particular, we consider spherical mirrors, which are simply mirrors in the shape of a small section of the surface of a sphere. A plane mirror is in fact a spherical mirror with an infinitely large *radius of curvature* and thus an approximately flat surface.

Making a Spherical Mirror

We start with the plane mirror of Fig. 34-8a, which faces leftward toward an object *O* that is shown and an observer that is not shown. We make a **concave mirror** by curving the mirror's surface so it is *concave* ("caved in") as in Fig. 34-8b. Curving the surface in this way changes several characteristics of the mirror and the image it produces of the object:



Bending it this way shifts the image closer.

1. The *center of curvature* *C* (the center of the sphere of which the mirror's surface is part) was infinitely far from the plane mirror; it is now closer but still in front of the concave mirror.
2. The *field of view*—the extent of the scene that is reflected to the observer—was wide; it is now smaller.
3. The image of the object was as far behind the plane mirror as the object was in front; the image is farther behind the concave mirror; that is, $|l|$ is greater.
4. The height of the image was equal to the height of the object; the height of the image is now greater. This feature is why many makeup mirrors and shaving mirrors are concave—they produce a larger image of a face.

We can make a **convex mirror** by curving a plane mirror so its surface is *convex* ("flexed out") as in Fig. 34-8c. Curving the surface in this way (1) moves the center of curvature *C* to *behind* the mirror and (2) *increases* the field of view. It also (3) moves the image of the object *closer* to the mirror and (4) *shrinks* it. Store surveillance mirrors are usually convex to take advantage of the increase in the field of view—more of the store can then be seen with a single mirror.

Fig. 34-8 (a) An object *O* forms a virtual image *I* in a plane mirror. (b) If the mirror is bent so that it becomes *concave*, the image moves farther away and becomes larger. (c) If the plane mirror is bent so that it becomes *convex*, the image moves closer and becomes smaller.

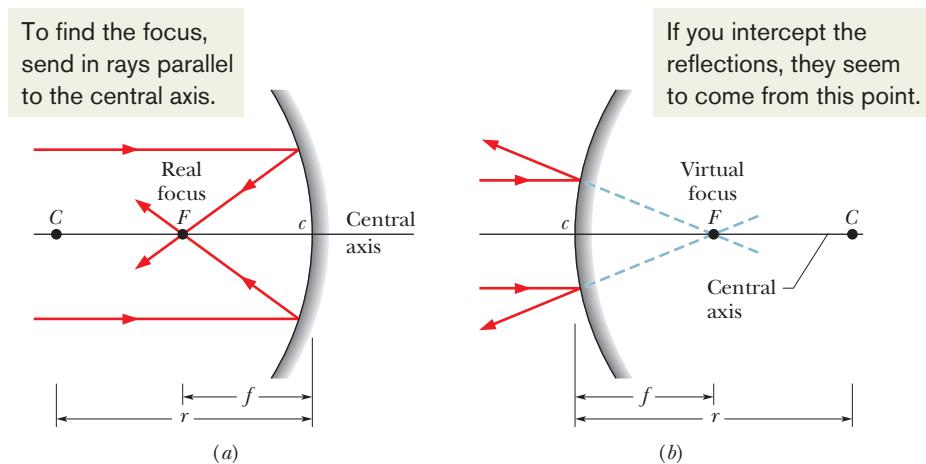


Fig. 34-9 (a) In a concave mirror, incident parallel light rays are brought to a real focus at F , on the same side of the mirror as the incident light rays. (b) In a convex mirror, incident parallel light rays seem to diverge from a virtual focus at F , on the side of the mirror opposite the light rays.

Focal Points of Spherical Mirrors

For a plane mirror, the magnitude of the image distance i is always equal to the object distance p . Before we can determine how these two distances are related for a spherical mirror, we must consider the reflection of light from an object O located an effectively infinite distance in front of a spherical mirror, on the mirror's *central axis*. That axis extends through the center of curvature C and the center c of the mirror. Because of the great distance between the object and the mirror, the light waves spreading from the object are plane waves when they reach the mirror along the central axis. This means that the rays representing the light waves are all parallel to the central axis when they reach the mirror.

When these parallel rays reach a concave mirror like that of Fig. 34-9a, those near the central axis are reflected through a common point F ; two of these reflected rays are shown in the figure. If we placed a (small) card at F , a point image of the infinitely distant object O would appear on the card. (This would occur for any infinitely distant object.) Point F is called the **focal point** (or **focus**) of the mirror, and its distance from the center of the mirror c is the **focal length** f of the mirror.

If we now substitute a convex mirror for the concave mirror, we find that the parallel rays are no longer reflected through a common point. Instead, they diverge as shown in Fig. 34-9b. However, if your eye intercepts some of the reflected light, you perceive the light as originating from a point source behind the mirror. This perceived source is located where extensions of the reflected rays pass through a common point (F in Fig. 34-9b). That point is the focal point (or focus) F of the convex mirror, and its distance from the mirror surface is the focal length f of the mirror. If we placed a card at this focal point, an image of object O would *not* appear on the card; so this focal point is not like that of a concave mirror.

To distinguish the actual focal point of a concave mirror from the perceived focal point of a convex mirror, the former is said to be a *real focal point* and the latter is said to be a *virtual focal point*. Moreover, the focal length f of a concave mirror is taken to be a positive quantity, and that of a convex mirror a negative quantity. For mirrors of both types, the focal length f is related to the radius of curvature r of the mirror by

$$f = \frac{1}{2}r \quad (\text{spherical mirror}), \quad (34-3)$$

where r is positive for a concave mirror and negative for a convex mirror.

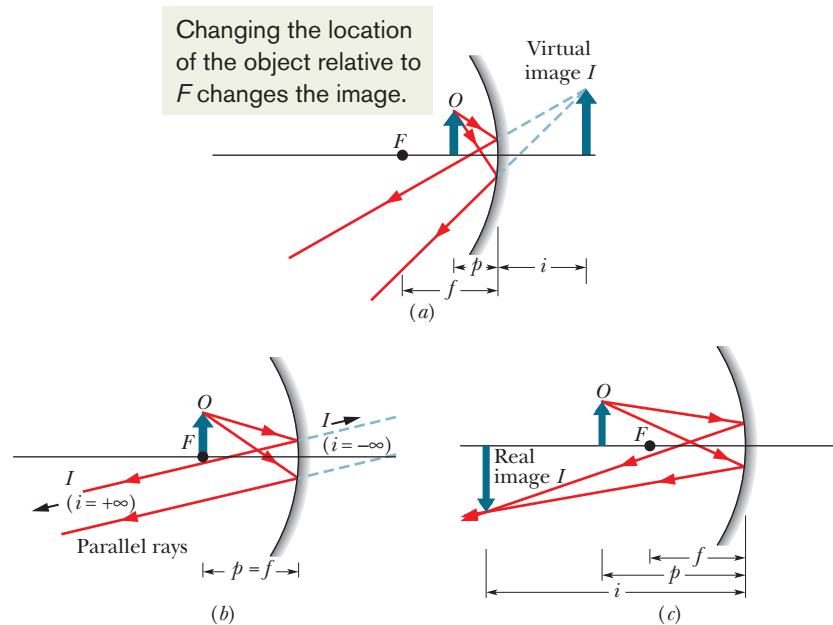


Fig. 34-10 (a) An object O inside the focal point of a concave mirror, and its virtual image I . (b) The object at the focal point F . (c) The object outside the focal point, and its real image I .

34-5 Images from Spherical Mirrors

With the focal point of a spherical mirror defined, we can find the relation between image distance i and object distance p for concave and convex spherical mirrors. We begin by placing the object O *inside the focal point* of the concave mirror—that is, between the mirror and its focal point F (Fig. 34-10a). An observer can then see a virtual image of O in the mirror: The image appears to be behind the mirror, and it has the same orientation as the object.

If we now move the object away from the mirror until it is at the focal point, the image moves farther and farther back from the mirror until, when the object is at the focal point, the image is at infinity (Fig. 34-10b). The image is then ambiguous and imperceptible because neither the rays reflected by the mirror nor the ray extensions behind the mirror cross to form an image of O .

If we next move the object *outside the focal point*—that is, farther away from the mirror than the focal point—the rays reflected by the mirror converge to form an *inverted* image of object O (Fig. 34-10c) in front of the mirror. That image moves in from infinity as we move the object farther outside F . If you were to hold a card at the position of the image, the image would show up on the card—the image is said to be *focused* on the card by the mirror. (The verb “focus,” which in this context means to produce an image, differs from the noun “focus,” which is another name for the focal point.) Because this image can actually appear on a surface, it is a real image—the rays actually intersect to create the image, regardless of whether an observer is present. The image distance i of a real image is a positive quantity, in contrast to that for a virtual image. We also see that



Real images form on the side of a mirror where the object is, and virtual images form on the opposite side.

34-5 IMAGES FROM SPHERICAL MIRRORS

931

As we shall prove in Section 34-9, when light rays from an object make only small angles with the central axis of a spherical mirror, a simple equation relates the object distance p , the image distance i , and the focal length f :

$$\frac{1}{p} + \frac{1}{i} = \frac{1}{f} \quad (\text{spherical mirror}). \quad (34-4)$$

We assume such small angles in figures such as Fig. 34-10, but for clarity the rays are drawn with exaggerated angles. With that assumption, Eq. 34-4 applies to any concave, convex, or plane mirror. For a convex or plane mirror, only a virtual image can be formed, regardless of the object's location on the central axis. As shown in the example of a convex mirror in Fig. 34-8c, the image is always on the opposite side of the mirror from the object and has the same orientation as the object.

The size of an object or image, as measured *perpendicular* to the mirror's central axis, is called the object or image *height*. Let h represent the height of the object, and h' the height of the image. Then the ratio h'/h is called the **lateral magnification** m produced by the mirror. However, by convention, the lateral magnification always includes a plus sign when the image orientation is that of the object and a minus sign when the image orientation is opposite that of the object. For this reason, we write the formula for m as

$$|m| = \frac{h'}{h} \quad (\text{lateral magnification}). \quad (34-5)$$

We shall soon prove that the lateral magnification can also be written as

$$m = -\frac{i}{p} \quad (\text{lateral magnification}). \quad (34-6)$$

For a plane mirror, for which $i = -p$, we have $m = +1$. The magnification of 1 means that the image is the same size as the object. The plus sign means that the image and the object have the same orientation. For the concave mirror of Fig. 34-10c, $m \approx -1.5$.

Equations 34-3 through 34-6 hold for all plane mirrors, concave spherical mirrors, and convex spherical mirrors. In addition to those equations, you have been asked to absorb a lot of information about these mirrors, and you should organize it for yourself by filling in Table 34-1. Under Image Location, note whether the image is on the *same* side of the mirror as the object or on the *opposite* side. Under Image Type, note whether the image is *real* or *virtual*. Under Image Orientation, note whether the image has the *same* orientation as the object or is *inverted*. Under Sign, give the sign of the quantity or fill in

Table 34-1

Your Organizing Table for Mirrors

Mirror Type	Object Location	Image			Sign		
		Location	Type	Orientation	of f	of r	of m
Plane	Anywhere						
Concave	Inside F						
	Outside F						
Convex	Anywhere						

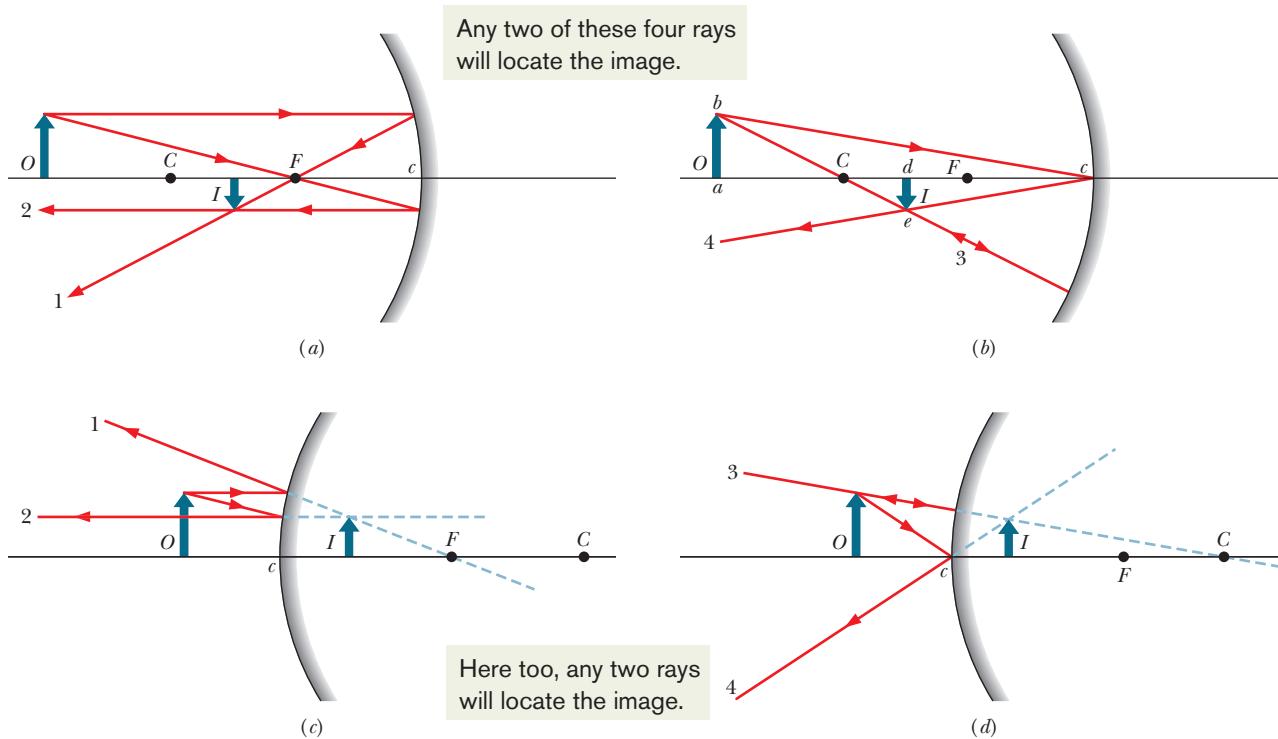


Fig. 34-11 (a, b) Four rays that may be drawn to find the image formed by a concave mirror. For the object position shown, the image is real, inverted, and smaller than the object. (c, d) Four similar rays for the case of a convex mirror. For a convex mirror, the image is always virtual, oriented like the object, and smaller than the object. [In (c), ray 2 is initially directed toward focal point F . In (d), ray 3 is initially directed toward center of curvature C .]

± if the sign is ambiguous. You will need this organization to tackle homework or a test.

Locating Images by Drawing Rays

Figures 34-11a and b show an object O in front of a concave mirror. We can graphically locate the image of any off-axis point of the object by drawing a *ray diagram* with any two of four special rays through the point:

1. A ray that is initially parallel to the central axis reflects through the focal point F (ray 1 in Fig. 34-11a).
2. A ray that reflects from the mirror after passing through the focal point emerges parallel to the central axis (ray 2 in Fig. 34-11a).
3. A ray that reflects from the mirror after passing through the center of curvature C returns along itself (ray 3 in Fig. 34-11b).
4. A ray that reflects from the mirror at point c is reflected symmetrically about that axis (ray 4 in Fig. 34-11b).

The image of the point is at the intersection of the two special rays you choose. The image of the object can then be found by locating the images of two or more of its off-axis points. You need to modify the descriptions of the rays slightly to apply them to convex mirrors, as in Figs. 34-11c and d.

Proof of Equation 34-6

We are now in a position to derive Eq. 34-6 ($m = -i/p$), the equation for the lateral magnification of an object reflected in a mirror. Consider ray 4 in Fig. 34-11b. It is reflected at point c so that the incident and reflected rays make equal angles with the axis of the mirror at that point.

The two right triangles abc and dec in the figure are similar (have the same set of angles); so we can write

$$\frac{de}{ab} = \frac{cd}{ca}.$$

The quantity on the left (apart from the question of sign) is the lateral magnification m produced by the mirror. Because we indicate an inverted image as a *negative* magnification, we symbolize this as $-m$. However, $cd = i$ and $ca = p$; so we have

$$m = -\frac{i}{p} \quad (\text{magnification}), \quad (34-7)$$

which is the relation we set out to prove.



CHECKPOINT 2

A Central American vampire bat, dozing on the central axis of a spherical mirror, is magnified by $m = -4$. Is its image (a) real or virtual, (b) inverted or of the same orientation as the bat, and (c) on the same side of the mirror as the bat or on the opposite side?

Sample Problem

Image produced by a spherical mirror

A tarantula of height h sits cautiously before a spherical mirror whose focal length has absolute value $|f| = 40$ cm. The image of the tarantula produced by the mirror has the same orientation as the tarantula and has height $h' = 0.20h$.

- (a) Is the image real or virtual, and is it on the same side of the mirror as the tarantula or the opposite side?

Reasoning: Because the image has the same orientation as the tarantula (the object), it must be virtual and on the opposite side of the mirror. (You can easily see this result if you have filled out Table 34-1.)

- (b) Is the mirror concave or convex, and what is its focal length f , sign included?

KEY IDEA

We *cannot* tell the type of mirror from the type of image because both types of mirror can produce virtual images. Similarly, we cannot tell the type of mirror from the sign of the focal length f , as obtained from Eq. 34-3 or Eq. 34-4, because we lack enough information to use either equation. However, we can make use of the magnification information.

Calculations: We know that the ratio of image height h' to object height h is 0.20. Thus, from Eq. 34-5 we have

$$|m| = \frac{h'}{h} = 0.20.$$

Because the object and image have the same orientation, we know that m must be positive: $m = +0.20$. Substituting this into Eq. 34-6 and solving for, say, i gives us

$$i = -0.20p,$$

which does not appear to be of help in finding f . However, it is helpful if we substitute it into Eq. 34-4. That equation gives us

$$\frac{1}{f} = \frac{1}{i} + \frac{1}{p} = \frac{1}{-0.20p} + \frac{1}{p} = \frac{1}{p}(-5 + 1),$$

from which we find

$$f = -p/4.$$

Now we have it: Because p is positive, f must be negative, which means that the mirror is convex with

$$f = -40 \text{ cm.} \quad (\text{Answer})$$



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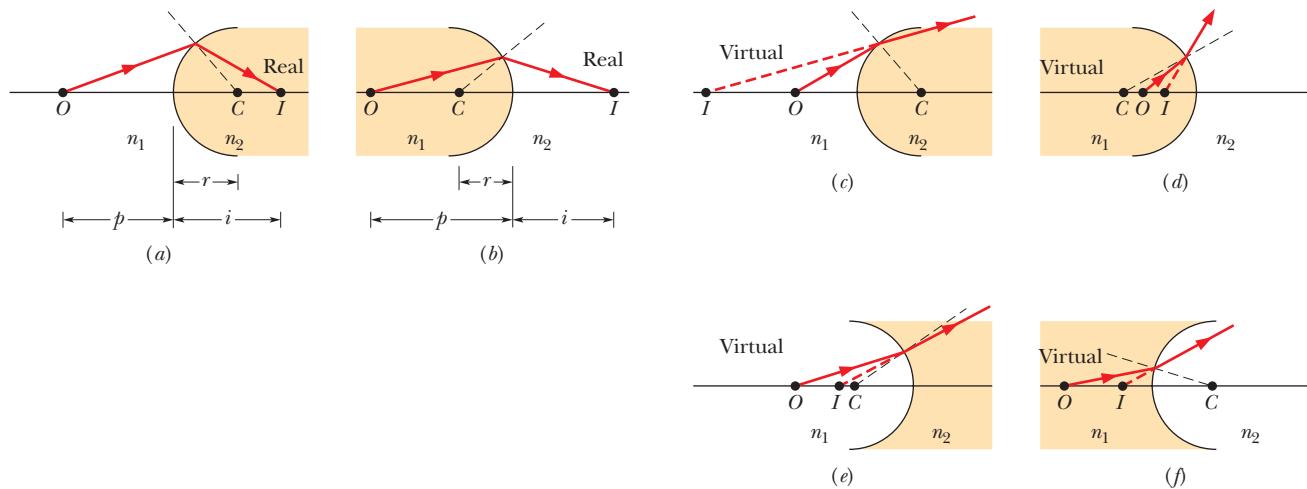


Fig. 34-12 Six possible ways in which an image can be formed by refraction through a spherical surface of radius r and center of curvature C . The surface separates a medium with index of refraction n_1 from a medium with index of refraction n_2 . The point object O is always in the medium with n_1 , to the left of the surface. The material with the lesser index of refraction is unshaded (think of it as being air, and the other material as being glass). Real images are formed in (a) and (b); virtual images are formed in the other four situations.

34-6 Spherical Refracting Surfaces

We now turn from images formed by reflection to images formed by refraction through surfaces of transparent materials, such as glass. We shall consider only spherical surfaces, with radius of curvature r and center of curvature C . The light will be emitted by a point object O in a medium with index of refraction n_1 ; it will refract through a spherical surface into a medium of index of refraction n_2 .

Our concern is whether the light rays, after refracting through the surface, form a real image (no observer necessary) or a virtual image (assuming that an observer intercepts the rays). The answer depends on the relative values of n_1 and n_2 and on the geometry of the situation.

Six possible results are shown in Fig. 34-12. In each part of the figure, the medium with the greater index of refraction is shaded, and object O is always in the medium with index of refraction n_1 , to the left of the refracting surface. In each part, a representative ray is shown refracting through the surface. (That ray and a ray along the central axis suffice to determine the position of the image in each case.)

At the point of refraction of each ray, the normal to the refracting surface is a radial line through the center of curvature C . Because of the refraction, the ray bends toward the normal if it is entering a medium of greater index of refraction and away from the normal if it is entering a medium of lesser index of refraction. If the bending sends the ray toward the central axis, that ray and others (undrawn) form a real image on that axis. If the bending sends the ray away from the central axis, the ray cannot form a real image; however, backward extensions of it and other refracted rays can form a virtual image, provided (as with mirrors) some of those rays are intercepted by an observer.

Real images I are formed (at image distance i) in parts *a* and *b* of Fig. 34-12, where the refraction directs the ray *toward* the central axis. Virtual images are formed in parts *c* and *d*, where the refraction directs the ray *away* from the central axis. Note, in these four parts, that real images are formed when the object is relatively far from the refracting surface and virtual images are formed when the object is nearer the refracting surface. In the final situations (Figs. 34-12*e* and *f*), refraction always directs the ray *away* from the central axis and virtual images are always formed, regardless of the object distance.



This insect has been entombed in amber for about 25 million years. Because we view the insect through a curved refracting surface, the location of the image we see does not coincide with the location of the insect (see Fig. 34-12*d*). (Dr. Paul A. Zahl/Photo Researchers)

34-6 SPHERICAL REFRACTING SURFACES

935

Note the following major difference from reflected images:



Real images form on the side of a refracting surface that is opposite the object, and virtual images form on the same side as the object.

In Section 34-9, we shall show that (for light rays making only small angles with the central axis)

$$\frac{n_1}{p} + \frac{n_2}{i} = \frac{n_2 - n_1}{r}. \quad (34-8)$$

Just as with mirrors, the object distance p is positive, and the image distance i is positive for a real image and negative for a virtual image. However, to keep all the signs correct in Eq. 34-8, we must use the following rule for the sign of the radius of curvature r :



When the object faces a convex refracting surface, the radius of curvature r is positive. When it faces a concave surface, r is negative.

Be careful: This is just the reverse of the sign convention we have for mirrors.

**CHECKPOINT 3**

A bee is hovering in front of the concave spherical refracting surface of a glass sculpture. (a) Which part of Fig. 34-12 is like this situation? (b) Is the image produced by the surface real or virtual, and (c) is it on the same side as the bee or the opposite side?

Sample Problem**Image produced by a refracting surface**

A Jurassic mosquito is discovered embedded in a chunk of amber, which has index of refraction 1.6. One surface of the amber is spherically convex with radius of curvature 3.0 mm (Fig. 34-13). The mosquito's head happens to be on the central axis of that surface and, when viewed along the axis, appears to be buried 5.0 mm into the amber. How deep is it really?

KEY IDEAS

The head appears to be 5.0 mm into the amber only because the light rays that the observer intercepts are bent by refraction at the convex amber surface. The image distance i differs from the object distance p according to Eq. 34-8. To use that equation to find the object distance, we first note:

1. Because the object (the head) and its image are on the same side of the refracting surface, the image must be virtual and so $i = -5.0$ mm.
2. Because the object is always taken to be in the medium of index of refraction n_1 , we must have $n_1 = 1.6$ and $n_2 = 1.0$.
3. Because the object faces a concave refracting surface, the radius of curvature r is negative, and so $r = -3.0$ mm.

Calculations: Making these substitutions in Eq. 34-8,

$$\begin{aligned} \frac{n_1}{p} + \frac{n_2}{i} &= \frac{n_2 - n_1}{r}, \\ \text{yields } \frac{1.6}{p} + \frac{1.0}{-5.0 \text{ mm}} &= \frac{1.0 - 1.6}{-3.0 \text{ mm}} \\ \text{and } p &= 4.0 \text{ mm.} \end{aligned} \quad (\text{Answer})$$

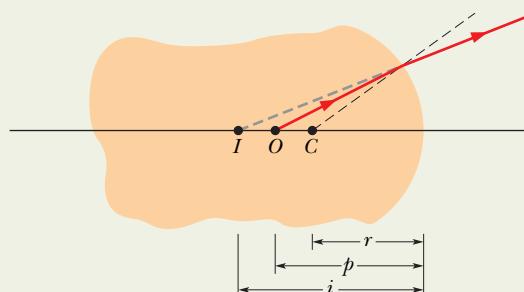


Fig. 34-13 A piece of amber with a mosquito from the Jurassic period, with the head buried at point O . The spherical refracting surface at the right end, with center of curvature C , provides an image I to an observer intercepting rays from the object at O .



Additional examples, video, and practice available at WileyPLUS

34-7 Thin Lenses

A **lens** is a transparent object with two refracting surfaces whose central axes coincide. The common central axis is the central axis of the lens. When a lens is surrounded by air, light refracts from the air into the lens, crosses through the lens, and then refracts back into the air. Each refraction can change the direction of travel of the light.

A lens that causes light rays initially parallel to the central axis to converge is (reasonably) called a **converging lens**. If, instead, it causes such rays to diverge, the lens is a **diverging lens**. When an object is placed in front of a lens of either type, light rays from the object that refract into and out of the lens can produce an image of the object.

We shall consider only the special case of a **thin lens**—that is, a lens in which the thickest part is thin relative to the object distance p , the image distance i , and the radii of curvature r_1 and r_2 of the two surfaces of the lens. We shall also consider only light rays that make small angles with the central axis (they are exaggerated in the figures here). In Section 34-9 we shall prove that for such rays, a thin lens has a focal length f . Moreover, i and p are related to each other by

$$\frac{1}{f} = \frac{1}{p} + \frac{1}{i} \quad (\text{thin lens}), \quad (34-9)$$

which is the same as we had for mirrors. We shall also prove that when a thin lens with index of refraction n is surrounded by air, this focal length f is given by

$$\frac{1}{f} = (n - 1) \left(\frac{1}{r_1} - \frac{1}{r_2} \right) \quad (\text{thin lens in air}), \quad (34-10)$$

which is often called the *lens maker's equation*. Here r_1 is the radius of curvature of the lens surface nearer the object and r_2 is that of the other surface. The signs of these radii are found with the rules in Section 34-6 for the radii of spherical refracting surfaces. If the lens is surrounded by some medium other than air (say, corn oil) with index of refraction n_{medium} , we replace n in Eq. 34-10 with n/n_{medium} . Keep in mind the basis of Eqs. 34-9 and 34-10:



A lens can produce an image of an object only because the lens can bend light rays, but it can bend light rays only if its index of refraction differs from that of the surrounding medium.

A fire is being started by focusing sunlight onto newspaper by means of a converging lens made of clear ice. The lens was made by melting both sides of a flat piece of ice into a convex shape in the shallow vessel (which has a curved bottom).

(Courtesy Matthew G. Wheeler)



Fig. 34-14 (a) Rays initially parallel to the central axis of a converging lens are made to converge to a real focal point F_2 by the lens. The lens is thinner than drawn, with a width like that of the vertical line through it. We shall consider all the bending of rays as occurring at this central line. (b) An enlargement of the top part of the lens of (a); normals to the surfaces are shown dashed. Note that both refractions bend the ray downward, toward the central axis. (c) The same initially parallel rays are made to diverge by a diverging lens. Extensions of the diverging rays pass through a virtual focal point F_2 . (d) An enlargement of the top part of the lens of (c). Note that both refractions bend the ray upward, away from the central axis.

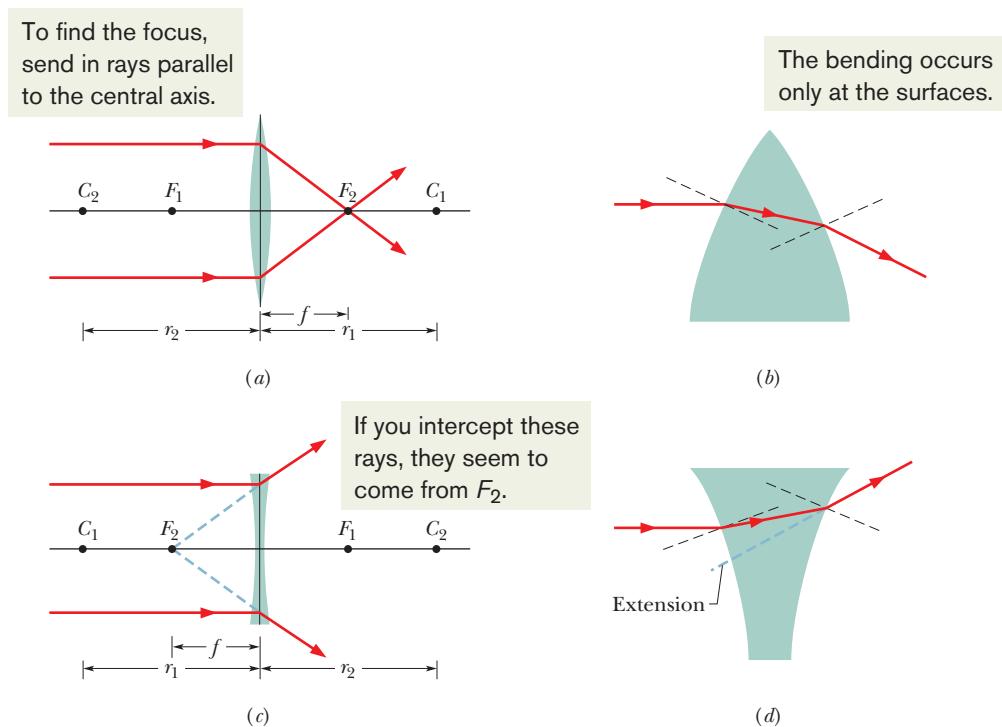


Figure 34-14a shows a thin lens with convex refracting surfaces, or *sides*. When rays that are parallel to the central axis of the lens are sent through the lens, they refract twice, as is shown enlarged in Fig. 34-14b. This double refraction causes the rays to converge and pass through a common point F_2 at a distance f from the center of the lens. Hence, this lens is a converging lens; further, a *real* focal point (or focus) exists at F_2 (because the rays really do pass through it), and the associated focal length is f . When rays parallel to the central axis are sent in the opposite direction through the lens, we find another real focal point at F_1 on the other side of the lens. For a thin lens, these two focal points are equidistant from the lens.

Because the focal points of a converging lens are real, we take the associated focal lengths f to be positive, just as we do with a real focus of a concave mirror. However, signs in optics can be tricky; so we had better check this in Eq. 34-10. The left side of that equation is positive if f is positive; how about the right side? We examine it term by term. Because the index of refraction n of glass or any other material is greater than 1, the term $(n - 1)$ must be positive. Because the source of the light (which is the object) is at the left and faces the convex left side of the lens, the radius of curvature r_1 of that side must be positive according to the sign rule for refracting surfaces. Similarly, because the object faces a concave right side of the lens, the radius of curvature r_2 of that side must be negative according to that rule. Thus, the term $(1/r_1 - 1/r_2)$ is positive, the whole right side of Eq. 34-10 is positive, and all the signs are consistent.

Figure 34-14c shows a thin lens with concave sides. When rays that are parallel to the central axis of the lens are sent through this lens, they refract twice, as is shown enlarged in Fig. 34-14d; these rays *diverge*, never passing through any common point, and so this lens is a diverging lens. However, extensions of the rays do pass through a common point F_2 at a distance f from the center of the lens. Hence, the lens has a *virtual* focal point at F_2 . (If your eye intercepts some of the diverging rays, you perceive a bright spot to be at F_2 , as if it is the source of the light.) Another virtual focus exists on the opposite side of the lens at F_1 , symmetrically placed if the lens is thin. Because the focal points of a diverging lens are virtual, we take the focal length f to be negative.

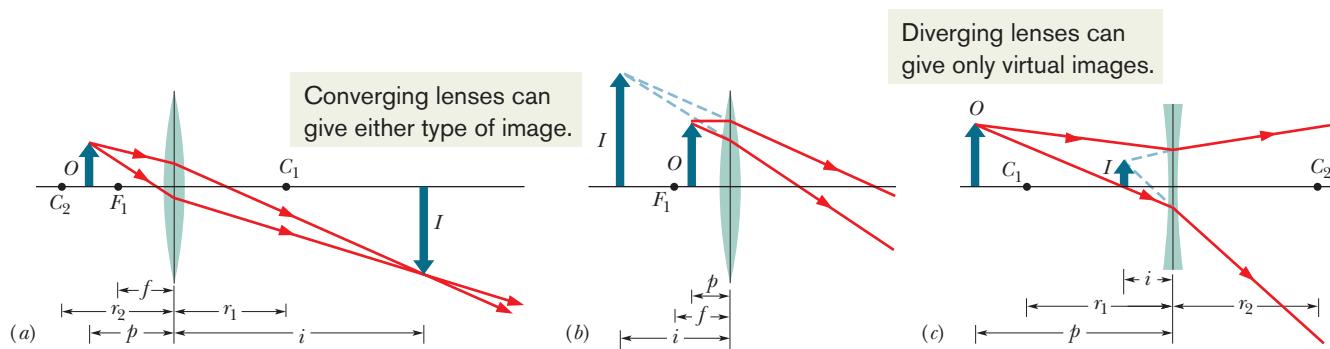


Fig. 34-15 (a) A real, inverted image I is formed by a converging lens when the object O is outside the focal point F_1 . (b) The image I is virtual and has the same orientation as O when O is inside the focal point. (c) A diverging lens forms a virtual image I , with the same orientation as the object O , whether O is inside or outside the focal point of the lens.

Images from Thin Lenses

We now consider the types of image formed by converging and diverging lenses. Figure 34-15a shows an object O outside the focal point F_1 of a converging lens. The two rays drawn in the figure show that the lens forms a real, inverted image I of the object on the side of the lens opposite the object.

When the object is placed inside the focal point F_1 , as in Fig. 34-15b, the lens forms a virtual image I on the same side of the lens as the object and with the same orientation. Hence, a converging lens can form either a real image or a virtual image, depending on whether the object is outside or inside the focal point, respectively.

Figure 34-15c shows an object O in front of a diverging lens. Regardless of the object distance (regardless of whether O is inside or outside the virtual focal point), this lens produces a virtual image that is on the same side of the lens as the object and has the same orientation.

As with mirrors, we take the image distance i to be positive when the image is real and negative when the image is virtual. However, the locations of real and virtual images from lenses are the reverse of those from mirrors:



Real images form on the side of a lens that is opposite the object, and virtual images form on the side where the object is.

The lateral magnification m produced by converging and diverging lenses is given by Eqs. 34-5 and 34-6, the same as for mirrors.

You have been asked to absorb a lot of information in this section, and you should organize it for yourself by filling in Table 34-2 for thin *symmetric lenses* (both

Table 34-2

Your Organizing Table for Thin Lenses

Lens Type	Object Location	Image			Sign		
		Location	Type	Orientation	off r	of r	of m
Converging	Inside F						
	Outside F						
Diverging	Anywhere						

sides are convex or both sides are concave). Under Image Location note whether the image is on the *same* side of the lens as the object or on the *opposite* side. Under Image Type note whether the image is *real* or *virtual*. Under Image Orientation note whether the image has the *same* orientation as the object or is *inverted*.

Locating Images of Extended Objects by Drawing Rays

Figure 34-16a shows an object O outside focal point F_1 of a converging lens. We can graphically locate the image of any off-axis point on such an object (such as the tip of the arrow in Fig. 34-16a) by drawing a ray diagram with any two of three special rays through the point. These special rays, chosen from all those that pass through the lens to form the image, are the following:

1. A ray that is initially parallel to the central axis of the lens will pass through focal point F_2 (ray 1 in Fig. 34-16a).
2. A ray that initially passes through focal point F_1 will emerge from the lens parallel to the central axis (ray 2 in Fig. 34-16a).
3. A ray that is initially directed toward the center of the lens will emerge from the lens with no change in its direction (ray 3 in Fig. 34-16a) because the ray encounters the two sides of the lens where they are almost parallel.

The image of the point is located where the rays intersect on the far side of the lens. The image of the object is found by locating the images of two or more of its points.

Figure 34-16b shows how the extensions of the three special rays can be used to locate the image of an object placed inside focal point F_1 of a converging lens. Note that the description of ray 2 requires modification (it is now a ray whose backward extension passes through F_1).

You need to modify the descriptions of rays 1 and 2 to use them to locate an image placed (anywhere) in front of a diverging lens. In Fig. 34-16c, for example, we find the point where ray 3 intersects the backward extensions of rays 1 and 2.

Two-Lens Systems

Here we consider an object sitting in front of a system of two lenses whose central axes coincide. Some of the possible two-lens systems are sketched in Fig. 34-17, but the figures are not drawn to scale. In each, the object sits to the left of lens 1 but can be inside or outside the focal point of the lens. Although tracing the light rays through any such two-lens system can be challenging, we can use the following simple two-step solution:

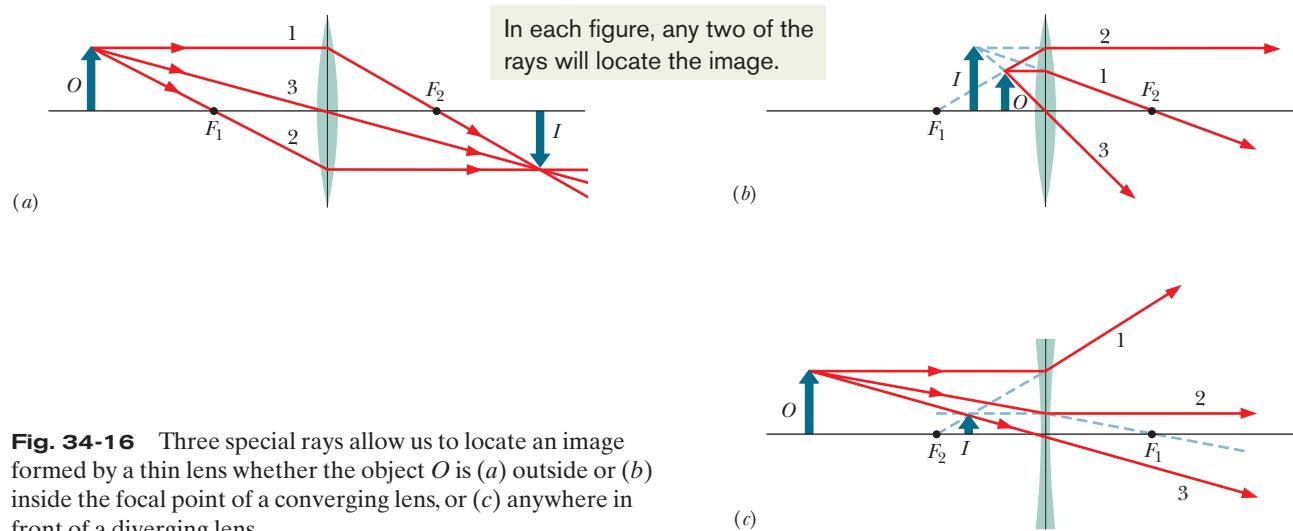


Fig. 34-16 Three special rays allow us to locate an image formed by a thin lens whether the object O is (a) outside or (b) inside the focal point of a converging lens, or (c) anywhere in front of a diverging lens.

Step 1 Neglecting lens 2, use Eq. 34-9 to locate the image I_1 produced by lens 1. Determine whether the image is on the left or right side of the lens, whether it is real or virtual, and whether it has the same orientation as the object. Roughly sketch I_1 . The top part of Fig. 34-17a gives an example.

Step 2 Neglecting lens 1, treat I_1 as though it is the *object* for lens 2. Use Eq. 34-9 to locate the image I_2 produced by lens 2. This is the final image of the system. Determine whether the image is on the left or right side of the lens, whether it is real or virtual, and whether it has the same orientation as the object for lens 2. Roughly sketch I_2 . The bottom part of Fig. 34-17a gives an example.

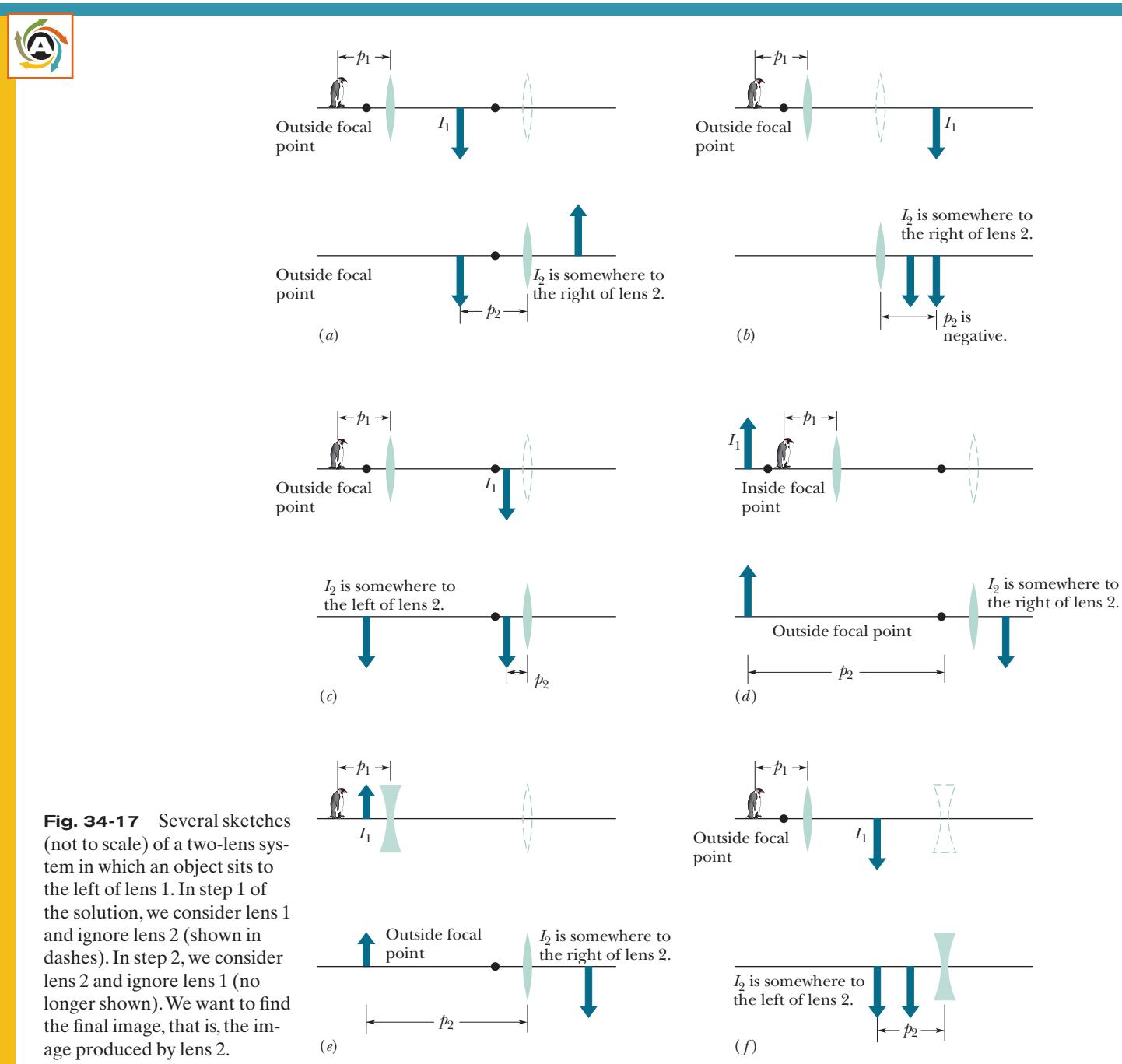


Fig. 34-17 Several sketches (not to scale) of a two-lens system in which an object sits to the left of lens 1. In step 1 of the solution, we consider lens 1 and ignore lens 2 (shown in dashes). In step 2, we consider lens 2 and ignore lens 1 (no longer shown). We want to find the final image, that is, the image produced by lens 2.

Thus we treat the two-lens system with two single-lens calculations, using the normal decisions and rules for a single lens. The only exception to the procedure occurs if I_1 lies to the right of lens 2 (past lens 2). We still treat it as the object for lens 2, but we take the object distance p_2 as a *negative* number when we use Eq. 34-9 to find I_2 . Then, as in our other examples, if the image distance i_2 is positive, the image is real and on the right side of the lens. An example is sketched in Fig. 34-17b.

This same step-by-step analysis can be applied for any number of lenses. It can also be applied if a mirror is substituted for lens 2. The *overall* (or *net*) lateral magnification M of a system of lenses (or lenses and a mirror) is the product of the individual lateral magnifications as given by Eq. 34-7 ($m = -i/p$). Thus, for a two-lens system, we have

$$M = m_1 m_2. \quad (34-11)$$

If M is positive, the final image has same the orientation as the object (the one in front of lens 1). If M is negative, the final image is inverted from the object. In the situation where p_2 is negative, such as in Fig. 34-17b, determining the orientation of the final image is probably easiest by examining the sign of M .



CHECKPOINT 4

A thin symmetric lens provides an image of a fingerprint with a magnification of +0.2 when the fingerprint is 1.0 cm farther from the lens than the focal point of the lens. What are the (a) type and (b) orientation of the image, and (c) what is the type of lens?

Sample Problem

Image produced by a thin symmetric lens

A praying mantis preys along the central axis of a thin symmetric lens, 20 cm from the lens. The lateral magnification of the mantis provided by the lens is $m = -0.25$, and the index of refraction of the lens material is 1.65.

- (a) Determine the type of image produced by the lens, the type of lens, whether the object (mantis) is inside or outside the focal point, on which side of the lens the image appears, and whether the image is inverted.

Reasoning: We can tell a lot about the lens and the image from the given value of m . From it and Eq. 34-6 ($m = -i/p$), we see that

$$i = -mp = 0.25p.$$

Even without finishing the calculation, we can answer the questions. Because p is positive, i here must be positive. That means we have a real image, which means we have a converging lens (the only lens that can produce a real image). The object must be outside the focal point (the only way a real image can be produced). Also, the image is inverted and on the side of the lens opposite the object. (That is how a converging lens makes a real image.)

- (b) What are the two radii of curvature of the lens?

KEY IDEAS

- Because the lens is symmetric, r_1 (for the surface nearer the object) and r_2 have the same magnitude r .
- Because the lens is a converging lens, the object faces a

convex surface on the nearer side and so $r_1 = +r$. Similarly, it faces a concave surface on the farther side; so $r_2 = -r$.

- We can relate these radii of curvature to the focal length f via the lens maker's equation, Eq. 34-10 (our only equation involving the radii of curvature of a lens).
- We can relate f to the object distance p and image distance i via Eq. 34-9.

Calculations: We know p , but we do not know i . Thus, our starting point is to finish the calculation for i in part (a); we obtain

$$i = (0.25)(20 \text{ cm}) = 5.0 \text{ cm}.$$

Now Eq. 34-9 gives us

$$\frac{1}{f} = \frac{1}{p} + \frac{1}{i} = \frac{1}{20 \text{ cm}} + \frac{1}{5.0 \text{ cm}},$$

from which we find $f = 4.0 \text{ cm}$.

Equation 34-10 then gives us

$$\frac{1}{f} = (n - 1) \left(\frac{1}{r_1} - \frac{1}{r_2} \right) = (n - 1) \left(\frac{1}{+r} - \frac{1}{-r} \right)$$

or, with known values inserted,

$$\frac{1}{4.0 \text{ cm}} = (1.65 - 1) \frac{2}{r},$$

which yields

$$r = (0.65)(2)(4.0 \text{ cm}) = 5.2 \text{ cm}. \quad (\text{Answer})$$



Additional examples, video, and practice available at WileyPLUS

Sample Problem

Image produced by a system of two thin lenses

Figure 34-18a shows a jalapeño seed O_1 that is placed in front of two thin symmetrical coaxial lenses 1 and 2, with focal lengths $f_1 = +24 \text{ cm}$ and $f_2 = +9.0 \text{ cm}$, respectively, and with lens separation $L = 10 \text{ cm}$. The seed is 6.0 cm from lens 1. Where does the system of two lenses produce an image of the seed?

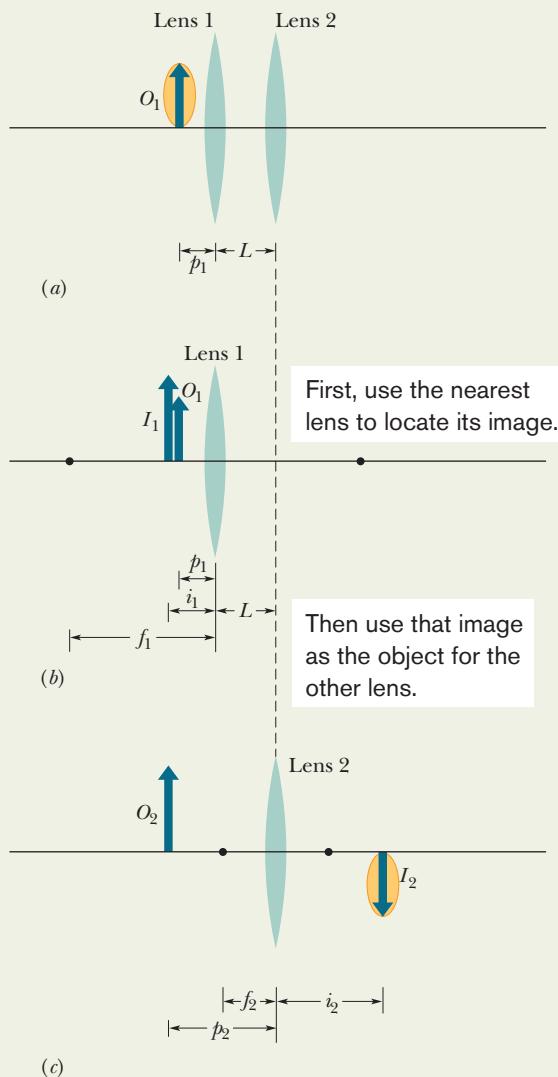


Fig. 34-18 (a) Seed O_1 is distance p_1 from a two-lens system with lens separation L . We use the arrow to orient the seed. (b) The image I_1 produced by lens 1 alone. (c) Image I_1 acts as object O_2 for lens 2 alone, which produces the final image I_2 .

KEY IDEA

We could locate the image produced by the system of lenses by tracing light rays from the seed through the two lenses. However, we can, instead, calculate the location of that image by working through the system in steps, lens by lens. We begin with the lens closer to the seed. The image we seek is the final one—that is, image I_2 produced by lens 2.

Lens 1: Ignoring lens 2, we locate the image I_1 produced by lens 1 by applying Eq. 34-9 to lens 1 alone:

$$\frac{1}{p_1} + \frac{1}{i_1} = \frac{1}{f_1}.$$

The object O_1 for lens 1 is the seed, which is 6.0 cm from the lens; thus, we substitute $p_1 = +6.0 \text{ cm}$. Also substituting the given value of f_1 , we then have

$$\frac{1}{+6.0 \text{ cm}} + \frac{1}{i_1} = \frac{1}{+24 \text{ cm}},$$

which yields $i_1 = -8.0 \text{ cm}$.

This tells us that image I_1 is 8.0 cm from lens 1 and virtual. (We could have guessed that it is virtual by noting that the seed is inside the focal point of lens 1, that is, between the lens and its focal point.) Because I_1 is virtual, it is on the same side of the lens as object O_1 and has the same orientation as the seed, as shown in Fig. 34-18b.

Lens 2: In the second step of our solution, we treat image I_1 as an object O_2 for the second lens and now ignore lens 1. We first note that this object O_2 is outside the focal point of lens 2. So the image I_2 produced by lens 2 must be real, inverted, and on the side of the lens opposite O_2 . Let us see.

The distance p_2 between this object O_2 and lens 2 is, from Fig. 34-18c,

$$p_2 = L + |i_1| = 10 \text{ cm} + 8.0 \text{ cm} = 18 \text{ cm}.$$

Then Eq. 34-9, now written for lens 2, yields

$$\frac{1}{+18 \text{ cm}} + \frac{1}{i_2} = \frac{1}{+9.0 \text{ cm}}.$$

Hence, $i_2 = +18 \text{ cm}$. (Answer)

The plus sign confirms our guess: Image I_2 produced by lens 2 is real, inverted, and on the side of lens 2 opposite O_2 , as shown in Fig. 34-18c. Thus, the image would appear on a card placed at its location.

34-8 Optical Instruments

The human eye is a remarkably effective organ, but its range can be extended in many ways by optical instruments such as eyeglasses, microscopes, and telescopes. Many such devices extend the scope of our vision beyond the visible range; satellite-borne infrared cameras and x-ray microscopes are just two examples.

The mirror and thin-lens formulas can be applied only as approximations to most sophisticated optical instruments. The lenses in typical laboratory microscopes are by no means “thin.” In most optical instruments the lenses are compound lenses; that is, they are made of several components, the interfaces rarely being exactly spherical. Now we discuss three optical instruments, assuming, for simplicity, that the thin-lens formulas apply.

Simple Magnifying Lens

The normal human eye can focus a sharp image of an object on the retina (at the rear of the eye) if the object is located anywhere from infinity to a certain point called the *near point* P_n . If you move the object closer to the eye than the near point, the perceived retinal image becomes fuzzy. The location of the near point normally varies with age. We have all heard about people who claim not to need glasses but read their newspapers at arm’s length; their near points are receding. To find your own near point, remove your glasses or contacts if you wear any, close one eye, and then bring this page closer to your open eye until it becomes indistinct. In what follows, we take the near point to be 25 cm from the eye, a bit more than the typical value for 20-year-olds.

Figure 34-19a shows an object O placed at the near point P_n of an eye. The size of the image of the object produced on the retina depends on the angle θ that the object occupies in the field of view from that eye. By moving the object closer to the eye, as in Fig. 34-19b, you can increase the angle and, hence, the possibility of distinguishing details of the object. However, because the object is

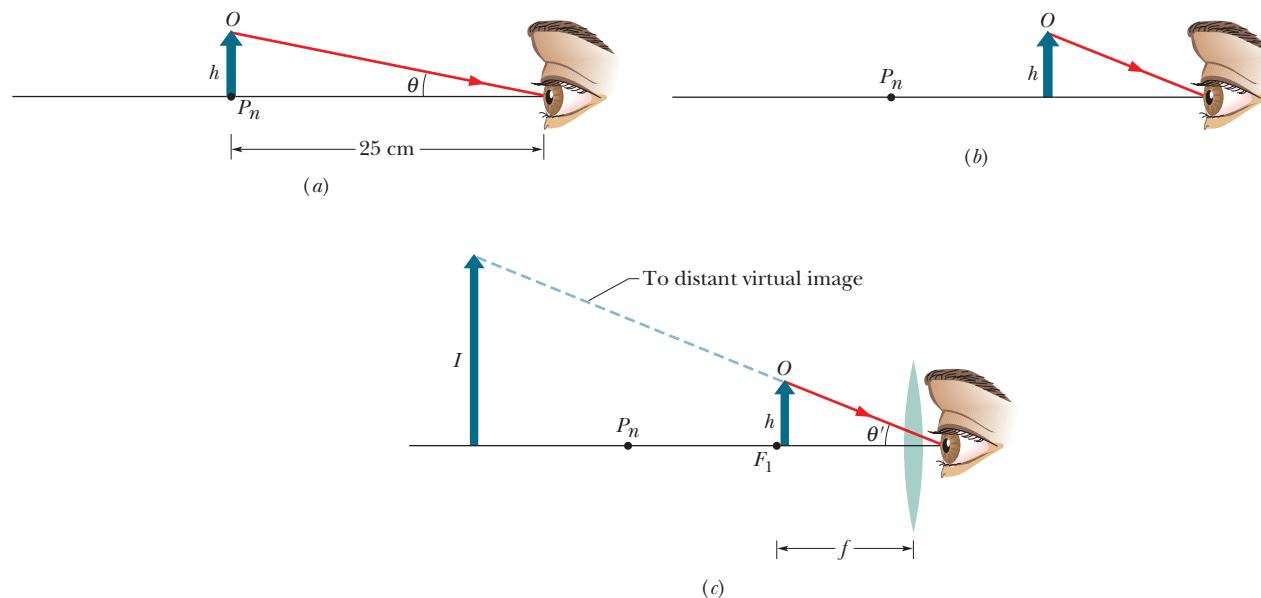


Fig. 34-19 (a) An object O of height h placed at the near point of a human eye occupies angle θ in the eye’s view. (b) The object is moved closer to increase the angle, but now the observer cannot bring the object into focus. (c) A converging lens is placed between the object and the eye, with the object just inside the focal point F_1 of the lens. The image produced by the lens is then far enough away to be focused by the eye, and the image occupies a larger angle θ' than object O does in (a).

then closer than the near point, it is no longer *in focus*; that is, the image is no longer clear.

You can restore the clarity by looking at O through a converging lens, placed so that O is just inside the focal point F_1 of the lens, which is at focal length f (Fig. 34-19c). What you then see is the virtual image of O produced by the lens. That image is farther away than the near point; thus, the eye can see it clearly.

Moreover, the angle θ' occupied by the virtual image is larger than the largest angle θ that the object alone can occupy and still be seen clearly. The *angular magnification* m_θ (not to be confused with lateral magnification m) of what is seen is

$$m_\theta = \theta'/\theta.$$

In words, the angular magnification of a simple magnifying lens is a comparison of the angle occupied by the image the lens produces with the angle occupied by the object when the object is moved to the near point of the viewer.

From Fig. 34-19, assuming that O is at the focal point of the lens, and approximating $\tan \theta$ as θ and $\tan \theta'$ as θ' for small angles, we have

$$\theta \approx h/25 \text{ cm} \quad \text{and} \quad \theta' \approx h/f.$$

We then find that

$$m_\theta \approx \frac{25 \text{ cm}}{f} \quad (\text{simple magnifier}). \quad (34-12)$$

Compound Microscope

Figure 34-20 shows a thin-lens version of a compound microscope. The instrument consists of an *objective* (the front lens) of focal length f_{ob} and an *eyepiece* (the lens near the eye) of focal length f_{ey} . It is used for viewing small objects that are very close to the objective.

The object O to be viewed is placed just outside the first focal point F_1 of the objective, close enough to F_1 that we can approximate its distance p from the lens as being f_{ob} . The separation between the lenses is then adjusted so that the enlarged, inverted, real image I produced by the objective is located just inside the first focal point F'_1 of the eyepiece. The *tube length* s shown in Fig. 34-20 is actually large relative to f_{ob} , and therefore we can approximate the distance i between the objective and the image I as being length s .

From Eq. 34-6, and using our approximations for p and i , we can write the lateral magnification produced by the objective as

$$m = -\frac{i}{p} = -\frac{s}{f_{ob}}. \quad (34-13)$$

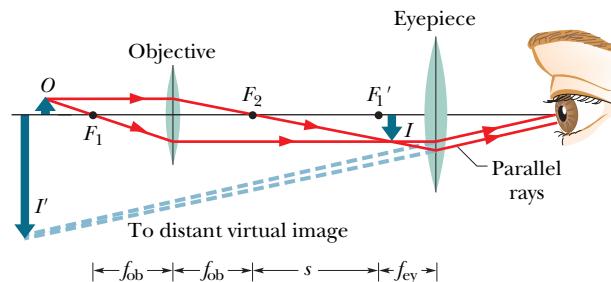


Fig. 34-20 A thin-lens representation of a compound microscope (not to scale). The objective produces a real image I of object O just inside the focal point F'_1 of the eyepiece. Image I then acts as an object for the eyepiece, which produces a virtual final image I' that is seen by the observer. The objective has focal length f_{ob} ; the eyepiece has focal length f_{ey} ; and s is the tube length.

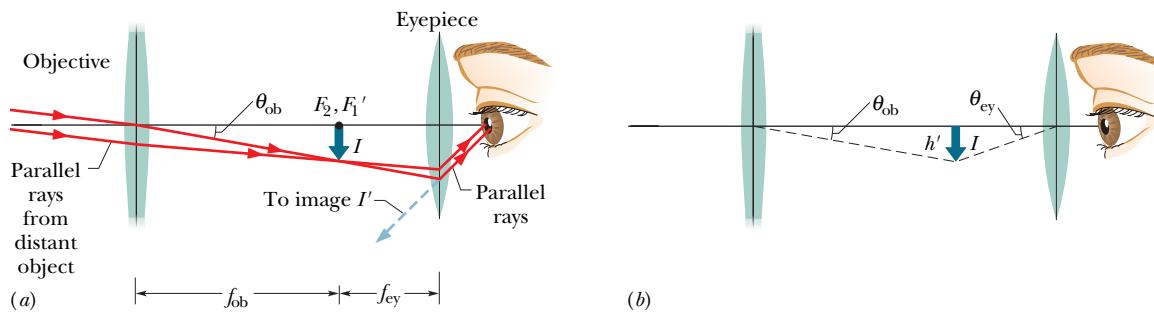


Fig. 34-21 (a) A thin-lens representation of a refracting telescope. From rays that are approximately parallel when they reach the objective, the objective produces a real image I of a distant source of light (the object). (One end of the object is assumed to lie on the central axis.) Image I , formed at the common focal points F_2 and F'_1 , acts as an object for the eyepiece, which produces a virtual final image I' at a great distance from the observer. The objective has focal length f_{ob} ; the eyepiece has focal length f_{ey} . (b) Image I has height h' and takes up angle θ_{ob} measured from the objective and angle θ_{ey} measured from the eyepiece.

Because the image I is located just inside the focal point F'_1 of the eyepiece, the eyepiece acts as a simple magnifying lens, and an observer sees a final (virtual, inverted) image I' through it. The overall magnification of the instrument is the product of the lateral magnification m produced by the objective, given by Eq. 34-13, and the angular magnification m_θ produced by the eyepiece, given by Eq. 34-12; that is,

$$M = mm_\theta = -\frac{s}{f_{\text{ob}}} \frac{25 \text{ cm}}{f_{\text{ey}}} \quad (\text{microscope}). \quad (34-14)$$

Refracting Telescope

Telescopes come in a variety of forms. The form we describe here is the simple refracting telescope that consists of an objective and an eyepiece; both are represented in Fig. 34-21 with simple lenses, although in practice, as is also true for most microscopes, each lens is actually a compound lens system.

The lens arrangements for telescopes and for microscopes are similar, but telescopes are designed to view large objects, such as galaxies, stars, and planets, at large distances, whereas microscopes are designed for just the opposite purpose. This difference requires that in the telescope of Fig. 34-21 the second focal point of the objective F_2 coincide with the first focal point of the eyepiece F'_1 , whereas in the microscope of Fig. 34-20 these points are separated by the tube length s .

In Fig. 34-21a, parallel rays from a distant object strike the objective, making an angle θ_{ob} with the telescope axis and forming a real, inverted image I at the common focal point F_2, F'_1 . This image I acts as an object for the eyepiece, through which an observer sees a distant (still inverted) virtual image I' . The rays defining the image make an angle θ_{ey} with the telescope axis.

The angular magnification m_θ of the telescope is $\theta_{\text{ey}}/\theta_{\text{ob}}$. From Fig. 34-21b, for rays close to the central axis, we can write $\theta_{\text{ob}} = h'/f_{\text{ob}}$ and $\theta_{\text{ey}} \approx h'/f_{\text{ey}}$, which gives us

$$m_\theta = -\frac{f_{\text{ob}}}{f_{\text{ey}}} \quad (\text{telescope}), \quad (34-15)$$

where the minus sign indicates that I' is inverted. In words, the angular magnification of a telescope is a comparison of the angle occupied by the image the telescope produces with the angle occupied by the distant object as seen without the telescope.

Magnification is only one of the design factors for an astronomical telescope and is indeed easily achieved. A good telescope needs *light-gathering power*, which determines how bright the image is. This is important for viewing faint objects such as distant galaxies and is accomplished by making the objective

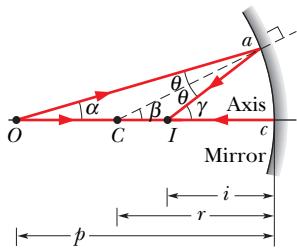


Fig. 34-22 A concave spherical mirror forms a real point image I by reflecting light rays from a point object O .

diameter as large as possible. A telescope also needs *resolving power*, which is the ability to distinguish between two distant objects (stars, say) whose angular separation is small. *Field of view* is another important design parameter. A telescope designed to look at galaxies (which occupy a tiny field of view) is much different from one designed to track meteors (which move over a wide field of view).

The telescope designer must also take into account the difference between real lenses and the ideal thin lenses we have discussed. A real lens with spherical surfaces does not form sharp images, a flaw called *spherical aberration*. Also, because refraction by the two surfaces of a real lens depends on wavelength, a real lens does not focus light of different wavelengths to the same point, a flaw called *chromatic aberration*.

This brief discussion by no means exhausts the design parameters of astronomical telescopes—many others are involved. We could make a similar listing for any other high-performance optical instrument.

34-9 Three Proofs

The Spherical Mirror Formula (Eq. 34-4)

Figure 34-22 shows a point object O placed on the central axis of a concave spherical mirror, outside its center of curvature C . A ray from O that makes an angle α with the axis intersects the axis at I after reflection from the mirror at a . A ray that leaves O along the axis is reflected back along itself at c and also passes through I . Thus, I is the image of O ; it is a *real* image because light actually passes through it. Let us find the image distance i .

A trigonometry theorem that is useful here tells us that an exterior angle of a triangle is equal to the sum of the two opposite interior angles. Applying this to triangles OaC and OaI in Fig. 34-22 yields

$$\beta = \alpha + \theta \quad \text{and} \quad \gamma = \alpha + 2\theta.$$

If we eliminate θ between these two equations, we find

$$\alpha + \gamma = 2\beta. \quad (34-16)$$

We can write angles α , β , and γ , in radian measure, as

$$\alpha \approx \frac{\widehat{ac}}{cO} = \frac{\widehat{ac}}{p}, \quad \beta = \frac{\widehat{ac}}{cC} = \frac{\widehat{ac}}{r},$$

$$\text{and} \quad \gamma \approx \frac{\widehat{ac}}{cI} = \frac{\widehat{ac}}{i}, \quad (34-17)$$

where the overhead symbol means “arc.” Only the equation for β is exact, because the center of curvature of \widehat{ac} is at C . However, the equations for α and γ are approximately correct if these angles are small enough (that is, for rays close to the central axis). Substituting Eqs. 34-17 into Eq. 34-16, using Eq. 34-3 to replace r with $2f$, and canceling \widehat{ac} lead exactly to Eq. 34-4, the relation that we set out to prove.

The Refracting Surface Formula (Eq. 34-8)

The incident ray from point object O in Fig. 34-23 that falls on point a of a spherical refracting surface is refracted there according to Eq. 33-40,

$$n_1 \sin \theta_1 = n_2 \sin \theta_2.$$

If α is small, θ_1 and θ_2 will also be small and we can replace the sines of these angles with the angles themselves. Thus, the equation above becomes

$$n_1 \theta_1 \approx n_2 \theta_2. \quad (34-18)$$

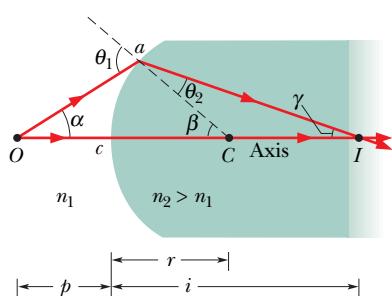


Fig. 34-23 A real point image I of a point object O is formed by refraction at a spherical convex surface between two media.

We again use the fact that an exterior angle of a triangle is equal to the sum of the two opposite interior angles. Applying this to triangles COa and ICa yields

$$\theta_1 = \alpha + \beta \quad \text{and} \quad \beta = \theta_2 + \gamma. \quad (34-19)$$

If we use Eqs. 34-19 to eliminate θ_1 and θ_2 from Eq. 34-18, we find

$$n_1\alpha + n_2\gamma = (n_2 - n_1)\beta. \quad (34-20)$$

In radian measure the angles α , β , and γ are

$$\alpha \approx \frac{\widehat{ac}}{p}; \quad \beta = \frac{\widehat{ac}}{r}; \quad \gamma \approx \frac{\widehat{ac}}{i}. \quad (34-21)$$

Only the second of these equations is exact. The other two are approximate because I and O are not the centers of circles of which \widehat{ac} is a part. However, for α small enough (for rays close to the axis), the inaccuracies in Eqs. 34-21 are small. Substituting Eqs. 34-21 into Eq. 34-20 leads directly to Eq. 34-8, as we wanted.

The Thin-Lens Formulas (Eqs. 34-9 and 34-10)

Our plan is to consider each lens surface as a separate refracting surface, and to use the image formed by the first surface as the object for the second.

We start with the thick glass "lens" of length L in Fig. 34-24a whose left and right refracting surfaces are ground to radii r' and r'' . A point object O' is placed near the left surface as shown. A ray leaving O' along the central axis is not deflected on entering or leaving the lens.

A second ray leaving O' at an angle α with the central axis intersects the left surface at point a' , is refracted, and intersects the second (right) surface at point a'' . The ray is again refracted and crosses the axis at I'' , which, being the intersection of two rays from O' , is the image of point O' , formed after refraction at two surfaces.

Figure 34-24b shows that the first (left) surface also forms a virtual image of O' at I' . To locate I' , we use Eq. 34-8,

$$\frac{n_1}{p} + \frac{n_2}{i} = \frac{n_2 - n_1}{r}.$$

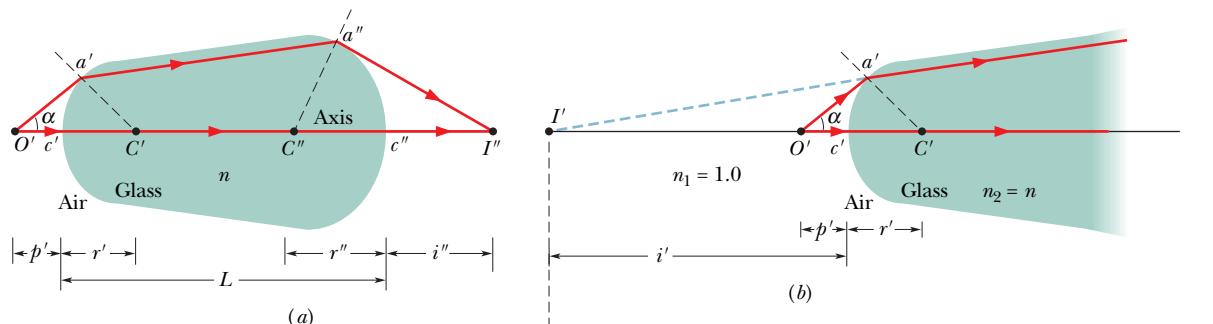
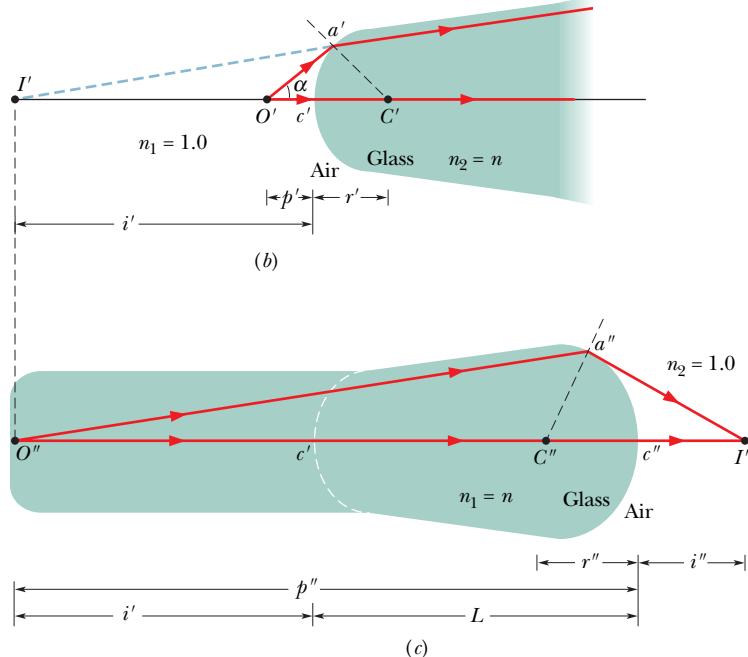


Fig. 34-24 (a) Two rays from point object O' form a real image I'' after refracting through two spherical surfaces of a lens. The object faces a convex surface at the left side of the lens and a concave surface at the right side. The ray traveling through points a' and a'' is actually close to the central axis through the lens. (b) The left side and (c) the right side of the lens in (a), shown separately.



Putting $n_1 = 1$ for air and $n_2 = n$ for lens glass and bearing in mind that the image distance is negative (that is, $i = -i'$ in Fig. 34-24b), we obtain

$$\frac{1}{p'} - \frac{n}{i'} = \frac{n-1}{r'}. \quad (34-22)$$

In this equation i' will be a positive number because we have already introduced the minus sign appropriate to a virtual image.

Figure 34-24c shows the second surface again. Unless an observer at point a'' were aware of the existence of the first surface, the observer would think that the light striking that point originated at point I' in Fig. 34-24b and that the region to the left of the surface was filled with glass as indicated. Thus, the (virtual) image I' formed by the first surface serves as a real object O'' for the second surface. The distance of this object from the second surface is

$$p'' = i' + L. \quad (34-23)$$

To apply Eq. 34-8 to the second surface, we must insert $n_1 = n$ and $n_2 = 1$ because the object now is effectively imbedded in glass. If we substitute with Eq. 34-23, then Eq. 34-8 becomes

$$\frac{n}{i' + L} + \frac{1}{i''} = \frac{1-n}{r''}. \quad (34-24)$$

Let us now assume that the thickness L of the “lens” in Fig. 34-24a is so small that we can neglect it in comparison with our other linear quantities (such as p' , i' , p'' , i'' , r' , and r''). In all that follows we make this *thin-lens approximation*. Putting $L = 0$ in Eq. 34-24 and rearranging the right side lead to

$$\frac{n}{i''} + \frac{1}{i''} = -\frac{n-1}{r''}. \quad (34-25)$$

Adding Eqs. 34-22 and 34-25 leads to

$$\frac{1}{p'} + \frac{1}{i''} = (n-1) \left(\frac{1}{r'} - \frac{1}{r''} \right).$$

Finally, calling the original object distance simply p and the final image distance simply i leads to

$$\frac{1}{p} + \frac{1}{i} = (n-1) \left(\frac{1}{r'} - \frac{1}{r''} \right), \quad (34-26)$$

which, with a small change in notation, is Eqs. 34-9 and 34-10.

REVIEW & SUMMARY

Real and Virtual Images An *image* is a reproduction of an object via light. If the image can form on a surface, it is a *real image* and can exist even if no observer is present. If the image requires the visual system of an observer, it is a *virtual image*.

Image Formation *Spherical mirrors, spherical refracting surfaces, and thin lenses* can form images of a source of light—the object—by redirecting rays emerging from the source. The image occurs where the redirected rays cross (forming a real image) or where backward extensions of those rays cross (forming a virtual image). If the rays are sufficiently close to the *central axis* through the spherical mirror, refracting surface, or thin lens, we have the following relations between the *object distance* p (which is positive) and the *image distance* i (which is positive for real images and negative for virtual images):

1. Spherical Mirror:

$$\frac{1}{p} + \frac{1}{i} = \frac{1}{f} = \frac{2}{r}, \quad (34-4, 34-3)$$

where f is the mirror’s focal length and r is its radius of curvature. A *plane mirror* is a special case for which $r \rightarrow \infty$, so that $p = -i$. Real images form on the side of a mirror where the object is located, and virtual images form on the opposite side.

2. Spherical Refracting Surface:

$$\frac{n_1}{p} + \frac{n_2}{i} = \frac{n_2 - n_1}{r} \quad (\text{single surface}), \quad (34-8)$$

where n_1 is the index of refraction of the material where the object is located, n_2 is the index of refraction of the material on the other side

of the refracting surface, and r is the radius of curvature of the surface. When the object faces a convex refracting surface, the radius r is positive. When it faces a concave surface, r is negative. Real images form on the side of a refracting surface that is opposite the object, and virtual images form on the same side as the object.

3. Thin Lens:

$$\frac{1}{p} + \frac{1}{i} = \frac{1}{f} = (n - 1) \left(\frac{1}{r_1} - \frac{1}{r_2} \right), \quad (34-9, 34-10)$$

where f is the lens's focal length, n is the index of refraction of the lens material, and r_1 and r_2 are the radii of curvature of the two sides of the lens, which are spherical surfaces. A convex lens surface that faces the object has a positive radius of curvature; a concave lens surface that faces the object has a negative radius of curvature. Real images form on the side of a lens that is opposite the object, and virtual images form on the same side as the object.

Lateral Magnification The *lateral magnification* m produced by a spherical mirror or a thin lens is

$$m = -\frac{i}{p}. \quad (34-6)$$

The magnitude of m is given by

$$|m| = \frac{h'}{h}, \quad (34-5)$$

**** View All Solutions Here ****

Q U E S T I O N S

- 1** Figure 34-25 shows a fish and a fish stalker in water. (a) Does the stalker see the fish in the general region of point a or point b ? (b) Does the fish see the (wild) eyes of the stalker in the general region of point c or point d ?

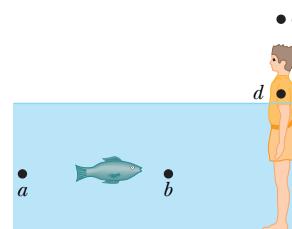


Fig. 34-25 Question 1.

- 2** In Fig. 34-26, stick figure O stands in front of a spherical mirror that is mounted within the boxed region; the central axis through the mirror is shown. The four stick figures I_1 to I_4 suggest general locations and orientations for the images that might be produced by the mirror. (The figures are only sketched in; neither their heights nor their distances from the mirror are drawn to scale.) (a) Which of the stick figures could not possibly represent images? Of the possible images, (b) which would be due to a concave mirror, (c) which would be due to a convex mirror, (d) which would be virtual, and (e) which would involve negative magnification?

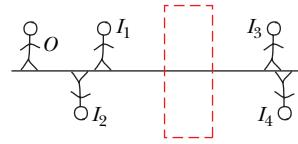


Fig. 34-26
Questions 2 and 10.

- 3** Figure 34-27 is an overhead view of a mirror maze based on floor sections that are equilateral triangles. Every wall within the maze is mirrored. If you stand at entrance x , (a) which of the maze monsters a , b , and c hiding in the maze can you see along

where h and h' are the heights (measured perpendicular to the central axis) of the object and image, respectively.

Optical Instruments Three optical instruments that extend human vision are:

- 1.** The *simple magnifying lens*, which produces an *angular magnification* m_θ given by

$$m_\theta = \frac{25 \text{ cm}}{f}, \quad (34-12)$$

where f is the focal length of the magnifying lens. The distance of 25 cm is a traditionally chosen value that is a bit more than the typical near point for someone 20 years old.

- 2.** The *compound microscope*, which produces an *overall magnification* M given by

$$M = mm_\theta = -\frac{s}{f_{ob}} \frac{25 \text{ cm}}{f_{ey}}, \quad (34-14)$$

where m is the lateral magnification produced by the objective, m_θ is the angular magnification produced by the eyepiece, s is the tube length, and f_{ob} and f_{ey} are the focal lengths of the objective and eyepiece, respectively.

- 3.** The *refracting telescope*, which produces an *angular magnification* m_θ given by

$$m_\theta = -\frac{f_{ob}}{f_{ey}}. \quad (34-15)$$

**** View All Solutions Here ****

the virtual hallways extending from entrance x ; (b) how many times does each visible monster appear in a hallway; and (c) what is at the far end of a hallway?

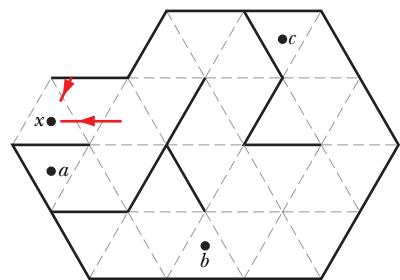


Fig. 34-27 Question 3.

- 4** A penguin waddles along the central axis of a concave mirror, from the focal point to an effectively infinite distance. (a) How does its image move? (b) Does the height of its image increase continuously, decrease continuously, or change in some more complicated manner?

- 5** When a *T. rex* pursues a jeep in the movie *Jurassic Park*, we see a reflected image of the *T. rex* via a side-view mirror, on which is printed the (then darkly humorous) warning: "Objects in mirror are closer than they appear." Is the mirror flat, convex, or concave?

- 6** An object is placed against the center of a concave mirror and then moved along the central axis until it is 5.0 m from the mirror. During the motion, the distance $|i|$ between the mirror and the image it produces is measured. The procedure is then repeated with a convex mirror and a plane mirror. Figure 34-28 gives the

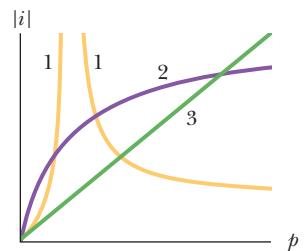


Fig. 34-28 Questions 6 and 8.

**** View All Solutions Here ****

results versus object distance p . Which curve corresponds to which mirror? (Curve 1 has two segments.)

- 7** The table details six variations of the basic arrangement of two thin lenses represented in Fig. 34-29. (The points labeled F_1 and F_2 are the focal points of lenses 1 and 2.) An object is distance p_1 to the left of lens 1, as in Fig. 34-18. (a) For which variations can we tell, *without calculation*, whether the final image (that due to lens 2) is to the left or right of lens 2 and whether it has the same orientation as the object? (b) For those “easy” variations, give the image location as “left” or “right” and the orientation as “same” or “inverted.”

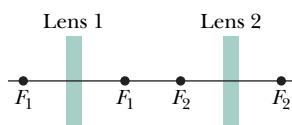


Fig. 34-29 Question 7.

Variation	Lens 1	Lens 2	
1	Converging	Converging	$p_1 < f_1 $
2	Converging	Converging	$p_1 > f_1 $
3	Diverging	Converging	$p_1 < f_1 $
4	Diverging	Converging	$p_1 > f_1 $
5	Diverging	Diverging	$p_1 < f_1 $
6	Diverging	Diverging	$p_1 > f_1 $

- 8** An object is placed against the center of a converging lens and then moved along the central axis until it is 5.0 m from the

lens. During the motion, the distance $|l|$ between the lens and the image it produces is measured. The procedure is then repeated with a diverging lens. Which of the curves in Fig. 34-28 best gives $|l|$ versus the object distance p for these lenses? (Curve 1 consists of two segments. Curve 3 is straight.)

- 9** Figure 34-30 shows four thin lenses, all of the same material, with sides that either are flat or have a radius of curvature of magnitude 10 cm. Without written calculation, rank the lenses according to the magnitude of the focal length, greatest first.

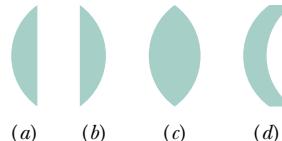


Fig. 34-30 Question 9.

- 10** In Fig. 34-26, stick figure O stands in front of a thin, symmetric lens that is mounted within the boxed region; the central axis through the lens is shown. The four stick figures I_1 to I_4 suggest general locations and orientations for the images that might be produced by the lens. (The figures are only sketched in; neither their height nor their distance from the lens is drawn to scale.) (a) Which of the stick figures could not possibly represent images? Of the possible images, (b) which would be due to a converging lens, (c) which would be due to a diverging lens, (d) which would be virtual, and (e) which would involve negative magnification?

PROBLEMS



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>



sec. 34-3 Plane Mirrors

- 1** You look through a camera toward an image of a hummingbird in a plane mirror. The camera is 4.30 m in front of the mirror. The bird is at camera level, 5.00 m to your right and 3.30 m from the mirror. What is the distance between the camera and the apparent position of the bird’s image in the mirror?

- 2** **ILW** A moth at about eye level is 10 cm in front of a plane mirror; you are behind the moth, 30 cm from the mirror. What is the distance between your eyes and the apparent position of the moth’s image in the mirror?

- 3** In Fig. 34-31, an isotropic point source of light S is positioned at distance d from a viewing screen A and the light intensity I_P at

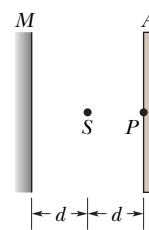


Fig. 34-31 Problem 3.

point P (level with S) is measured. Then a plane mirror M is placed behind S at distance d . By how much is I_P multiplied by the presence of the mirror?

- 4** Figure 34-32 shows an overhead view of a corridor with a plane mirror M mounted at one end. A burglar B sneaks along the corridor directly toward the center of the mirror. If $d = 3.0$ m, how far from the mirror will she be when the security guard S can first see her in the mirror?

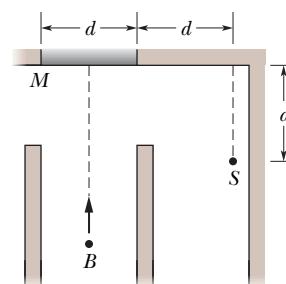


Fig. 34-32 Problem 4.

- 5** **SSM** **WWW** Figure 34-33 shows a small lightbulb suspended at distance $d_1 = 250$ cm above the surface of the water in a swimming

pool where the water depth is $d_2 = 200 \text{ cm}$. The bottom of the pool is a large mirror. How far below the mirror surface is the image of the bulb? (Hint: Assume that the rays are close to a vertical axis through the bulb, and use the small-angle approximation in which $\sin \theta \approx \tan \theta \approx \theta$.)

sec. 34-5 Images from Spherical Mirrors

•6 An object is moved along the central axis of a spherical mirror while the lateral magnification m of it is measured. Figure 34-34 gives m versus object distance p for the range $p_a = 2.0 \text{ cm}$ to $p_b = 8.0 \text{ cm}$. What is m for $p = 14.0 \text{ cm}$?

•7 A concave shaving mirror has a radius of curvature of 35.0 cm . It is positioned so that the (upright) image of a man's face is 2.50 times the size of the face. How far is the mirror from the face?

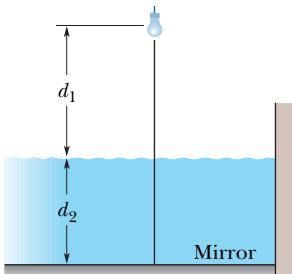


Fig. 34-33 Problem 5.

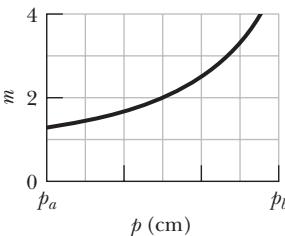


Fig. 34-34 Problem 6.

•8 An object is placed against the center of a spherical mirror and then moved 70 cm from it along the central axis as the image distance i is measured. Figure 34-35 gives i versus object distance p out to $p_s = 40 \text{ cm}$. What is i for $p = 70 \text{ cm}$?

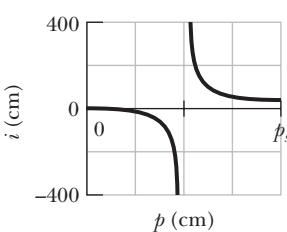


Fig. 34-35 Problem 8.

•9 through 16 **SSM** 9, 11, 13 **GO** 12 *Spherical mirrors.* Object O stands on the central axis of a spherical mirror. For this situation, each problem in Table 34-3 gives object distance p_s (centimeters), the type of mirror, and then the distance (centimeters, without proper sign) between the focal point and the mirror. Find (a) the radius of curvature r (including sign), (b) the image distance i , and (c) the lateral magnification m . Also, determine whether the image is (d) real (R) or virtual (V), (e) inverted (I) from object O or noninverted (NI), and (f) on the same side of the mirror as O or on the opposite side.

•17 through 29 **SSM** 23, 29 **GO** 22 *More mirrors.* Object O stands on the central axis of a spherical or plane mirror. For this situation, each problem in Table 34-4 refers to (a) the type of mirror,

Table 34-3

Problems 9 through 16: Spherical Mirrors. See the setup for these problems.

	p	Mirror	(a) r	(b) i	(c) m	(d) R/V	(e) I/NI	(f) Side
9	+18	Concave, 12						
10	+15	Concave, 10						
11	+8.0	Convex, 10						
12	+24	Concave, 36						
13	+12	Concave, 18						
14	+22	Convex, 35						
15	+10	Convex, 8.0						
16	+17	Convex, 14						

Table 34-4

Problems 17 through 29: More Mirrors. See the setup for these problems.

	(a) Type	(b) f	(c) r	(d) p	(e) i	(f) m	(g) R/V	(h) I/NI	(i) Side
17	Concave	20		+10					
18				+24					
19			-40		-10		0.50		I
20				+40			-0.70		
21		+20		+30					
22		20					+0.10		
23		30					+0.20		
24				+60			-0.50		
25				+30			0.40		I
26		20		+60					Same
27		-30			-15				
28				+10			+1.0		
29	Convex		40		4.0				

(b) the focal distance f , (c) the radius of curvature r , (d) the object distance p , (e) the image distance i , and (f) the lateral magnification m . (All distances are in centimeters.) It also refers to whether (g) the image is real (R) or virtual (V), (h) inverted (I) or noninverted (NI) from O , and (i) on the same side of the mirror as object O or on the opposite side. Fill in the missing information. Where only a sign is missing, answer with the sign.

- 30 Figure 34-36 gives the lateral magnification m of an object versus the object distance p from a spherical mirror as the object is moved along the mirror's central axis through a range of values for p . The horizontal scale is set by $p_s = 10.0$ cm. What is the magnification of the object when the object is 21 cm from the mirror?

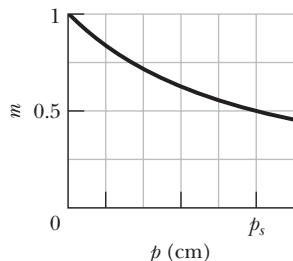


Fig. 34-36 Problem 30.

- 31 (a) A luminous point is moving at speed v_o toward a spherical mirror with radius of curvature r , along the central axis of the mirror. Show that the image of this point is moving at speed

$$v_I = -\left(\frac{r}{2p - r}\right)^2 v_o,$$

where p is the distance of the luminous point from the mirror at any given time. Now assume the mirror is concave, with $r = 15$ cm, and let $v_o = 5.0$ cm/s. Find v_I when (b) $p = 30$ cm (far outside the focal point), (c) $p = 8.0$ cm (just outside the focal point), and (d) $p = 10$ mm (very near the mirror).

sec. 34-6 Spherical Refracting Surfaces

- 32 through 38 **SSM** 33, 35 **GO** 37 *Spherical refracting surfaces.* An object O stands on the central axis of a spherical refracting surface. For this situation, each problem in Table 34-5 refers to the index of refraction n_1 where the object is located, (a) the index of refraction n_2 on the other side of the refracting surface, (b) the object distance p , (c) the radius of curvature r of the surface, and (d) the image distance i . (All distances are in centimeters.) Fill in the missing information, including whether the image is (e) real (R) or virtual (V) and (f) on the same side of the surface as object O or on the opposite side.

- 39 In Fig. 34-37, a beam of parallel light rays from a laser is incident

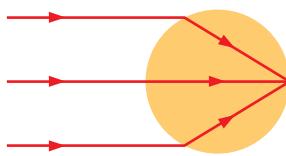


Fig. 34-37 Problem 39.

on a solid transparent sphere of index of refraction n . (a) If a point image is produced at the back of the sphere, what is the index of refraction of the sphere? (b) What index of refraction, if any, will produce a point image at the center of the sphere?

- 40 A glass sphere has radius $R = 5.0$ cm and index of refraction 1.6. A paperweight is constructed by slicing through the sphere along a plane that is 2.0 cm from the center of the sphere, leaving height $h = 3.0$ cm. The paperweight is placed on a table and viewed from directly above by an observer who is distance $d = 8.0$ cm from the tabletop (Fig. 34-38). When viewed through the paperweight, how far away does the tabletop appear to be to the observer?

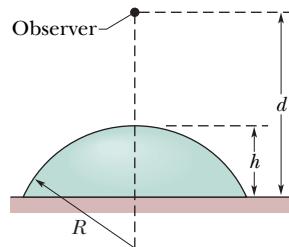


Fig. 34-38 Problem 40.

sec. 34-7 Thin Lenses

- 41 A lens is made of glass having an index of refraction of 1.5. One side of the lens is flat, and the other is convex with a radius of curvature of 20 cm. (a) Find the focal length of the lens. (b) If an object is placed 40 cm in front of the lens, where is the image?

- 42 Figure 34-39 gives the lateral magnification m of an object versus the object distance p from a lens as the object is moved along the central axis of the lens through a range of values for p out to $p_s = 20.0$ cm. What is the magnification of the object when the object is 35 cm from the lens?

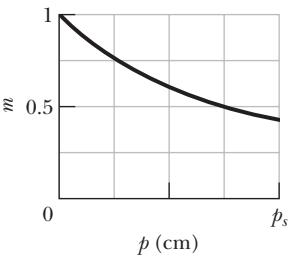


Fig. 34-39 Problem 42.

- 43 A movie camera with a (single) lens of focal length 75 mm takes a picture of a person standing 27 m away. If the person is 180 cm tall, what is the height of the image on the film?

- 44 An object is placed against the center of a thin lens and then moved away from it along the central axis as the image distance i is measured. Figure 34-40 gives i versus object distance p out to $p_s = 60$ cm. What is the image distance when $p = 100$ cm?

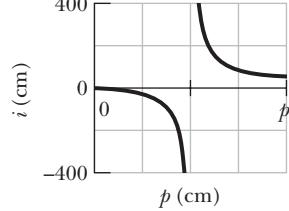


Fig. 34-40 Problem 44.

Table 34-5

Problems 32 through 38: Spherical Refracting Surfaces.
See the setup for these problems.

	(a)	(b)	(c)	(d)	(e)	(f)
n_1	n_2	p	r	i	R/V	Side
32	1.0	1.5	+10	+30		
33	1.0	1.5	+10		-13	
34	1.5		+100	-30	+600	
35	1.5	1.0	+70	+30		
36	1.5	1.0		-30	-7.5	
37	1.5	1.0	+10		-6.0	
38	1.0	1.5		+30	+600	

Table 34-6

Problems 50 through 57: Thin Lenses. See the setup for these problems.

	<i>p</i>	Lens	(a) <i>i</i>	(b) <i>m</i>	(c) R/V	(d) I/NI	(e) Side
50	+16	C, 4.0					
51	+12	C, 16					
52	+25	C, 35					
53	+8.0	D, 12					
54	+10	D, 6.0					
55	+22	D, 14					
56	+12	D, 31					
57	+45	C, 20					

- 45 You produce an image of the Sun on a screen, using a thin lens whose focal length is 20.0 cm. What is the diameter of the image? (See Appendix C for needed data on the Sun.)

- 46 An object is placed against the center of a thin lens and then moved 70 cm from it along the central axis as the image distance *i* is measured. Figure 34-41 gives *i* versus object distance *p* out to $p_s = 40$ cm. What is the image distance when *p* = 70 cm?

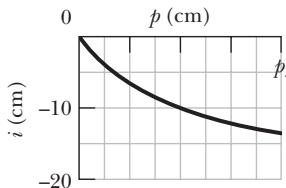


Fig. 34-41 Problem 46.

- 47 SSM WWW A double-convex lens is to be made of glass with an index of refraction of 1.5. One surface is to have twice the radius of curvature of the other and the focal length is to be 60 mm. What is the (a) smaller and (b) larger radius?

- 48 An object is moved along the central axis of a thin lens while the lateral magnification *m* is measured. Figure 34-42 gives *m* versus object distance *p* out to $p_s = 8.0$ cm. What is the magnification of the object when the object is 14.0 cm from the lens?

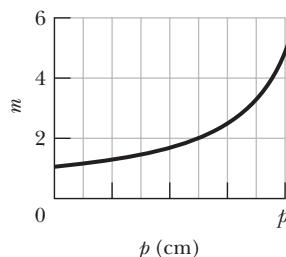


Fig. 34-42 Problem 48.

- 49 SSM An illuminated slide is held 44 cm from a screen. How far from the slide must a lens of focal length 11 cm be placed (between the slide and the screen) to form an image of the slide's picture on the screen?

- 50 through 57 SSM 53, 57 GO 55 Thin lenses. Object *O* stands on the central axis of a thin symmetric lens. For this situation, each problem in Table 34-6 gives object distance *p* (centimeters), the type of lens (C stands for converging and D for diverging), and then the distance (centimeters, without proper sign) between a focal point and the lens. Find (a) the image distance *i* and (b) the lateral magnification *m* of the object, including signs. Also, determine whether the image is (c) real (R) or virtual (V), (d) inverted (I) from object *O* or noninverted (NI), and (e) on the same side of the lens as object *O* or on the opposite side.

- 58 through 67 SSM 59 GO 66 Lenses with given radii. Object *O* stands in front of a thin lens, on the central axis. For this situation, each problem in Table 34-7 gives object distance *p*, index of refraction *n* of the lens, radius *r*₁ of the nearer lens surface, and radius *r*₂ of the farther lens surface. (All distances are in centimeters.) Find (a) the image distance *i* and (b) the lateral magnification *m* of the object, including signs. Also, determine whether the image is (c) real (R) or virtual (V), (d) inverted (I) from object *O* or noninverted (NI), and (e) on the same side of the lens as object *O* or on the opposite side.

Table 34-7

Problems 58 through 67: Lenses with Given Radii. See the setup for these problems.

	<i>p</i>	<i>n</i>	<i>r</i> ₁	<i>r</i> ₂	(a) <i>i</i>	(b) <i>m</i>	(c) R/V	(d) I/NI	(e) Side
58	+29	1.65	+35	∞					
59	+75	1.55	+30	-42					
60	+6.0	1.70	+10	-12					
61	+24	1.50	-15	-25					
62	+10	1.50	+30	-30					
63	+35	1.70	+42	+33					
64	+10	1.50	-30	-60					
65	+10	1.50	-30	+30					
66	+18	1.60	-27	+24					
67	+60	1.50	+35	-35					

Table 34-8

Problems 69 through 79: More Lenses. See the setup for these problems.

	(a) Type	(b) f	(c) p	(d) i	(e) m	(f) R/V	(g) I/NI	(h) Side
69		+10	+5.0					
70		20	+8.0		<1.0		NI	
71			+16		+0.25			
72			+16		-0.25			
73			+10		-0.50			
74	C	10	+20					
75		10	+5.0		<1.0		Same	
76		10	+5.0		>1.0			
77			+16		+1.25			
78			+10		0.50		NI	
79		20	+8.0		>1.0			

- 68 In Fig. 34-43, a real inverted image I of an object O is formed by a certain lens (not shown); the object-image separation is $d = 40.0\text{ cm}$, measured along the central axis of the lens. The image is just half the size of the object. (a) What kind of lens must be used to produce this image? (b) How far from the object must the lens be placed? (c) What is the focal length of the lens?

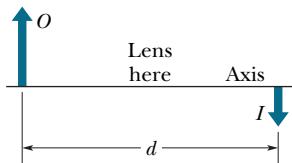


Fig. 34-43 Problem 68.

- 69 through 79 SSM 75, 77 GO 78 More lenses. Object O stands on the central axis of a thin symmetric lens. For this situation, each problem in Table 34-8 refers to (a) the lens type, converging (C) or diverging (D), (b) the focal distance f , (c) the object distance p , (d) the image distance i , and (e) the lateral

magnification m . (All distances are in centimeters.) It also refers to whether (f) the image is real (R) or virtual (V), (g) inverted (I) or noninverted (NI) from O , and (h) on the *same* side of the lens as O or on the *opposite* side. Fill in the missing information, including the value of m when only an inequality is given. Where only a sign is missing, answer with the sign.

- 80 through 87 SSM WWW 83 GO 82, 85 Two-lens systems. In Fig. 34-44, stick figure O (the object) stands on the common central axis of two thin, symmetric lenses, which are mounted in the boxed regions. Lens 1 is mounted within the boxed region closer to O , which is at object distance p_1 . Lens 2 is mounted within the farther boxed region, at distance d . Each problem in Table 34-9 refers to a different combination of lenses and different values for distances, which are given in centimeters. The type of lens is indicated by C for converging and D for diverging; the number after C or D is the distance between a lens and either

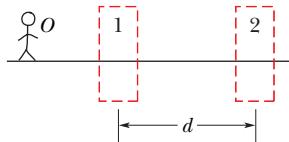


Fig. 34-44 Problems 80 through 87.

Table 34-9

Problems 80 through 87: Two-Lens Systems. See the setup for these problems.

	p_1	Lens 1	d	Lens 2	(a) i_2	(b) M	(c) R/V	(d) I/NI	(e) Side
80	+10	C, 15	10	C, 8.0					
81	+12	C, 8.0	32	C, 6.0					
82	+8.0	D, 6.0	12	C, 6.0					
83	+20	C, 9.0	8.0	C, 5.0					
84	+15	C, 12	67	C, 10					
85	+4.0	C, 6.0	8.0	D, 6.0					
86	+12	C, 8.0	30	D, 8.0					
87	+20	D, 12	10	D, 8.0					

of its focal points (the proper sign of the focal distance is not indicated).

Find (a) the image distance i_2 for the image produced by lens 2 (the final image produced by the system) and (b) the overall lateral magnification M for the system, including signs. Also, determine whether the final image is (c) real (R) or virtual (V), (d) inverted (I) from object O or noninverted (NI), and (e) on the same side of lens 2 as object O or on the opposite side.

sec. 34-8 Optical Instruments

•88 If the angular magnification of an astronomical telescope is 36 and the diameter of the objective is 75 mm, what is the minimum diameter of the eyepiece required to collect all the light entering the objective from a distant point source on the telescope axis?

•89 SSM In a microscope of the type shown in Fig. 34-20, the focal length of the objective is 4.00 cm, and that of the eyepiece is 8.00 cm. The distance between the lenses is 25.0 cm. (a) What is the tube length s ? (b) If image I in Fig. 34-20 is to be just inside focal point F'_1 , how far from the objective should the object be? What then are (c) the lateral magnification m of the objective, (d) the angular magnification m_θ of the eyepiece, and (e) the overall magnification M of the microscope?

•90 Figure 34-45a shows the basic structure of a camera. A lens can be moved forward or back to produce an image on film at the back of the camera. For a certain camera, with the distance i between the lens and the film set at $f = 5.0$ cm, parallel light rays from a very distant object O converge to a point image on the film, as shown. The object is now brought closer, to a distance of $p = 100$ cm, and the lens-film distance is adjusted so that an inverted real image forms on the film (Fig. 34-45b). (a) What is the lens-film distance i now? (b) By how much was distance i changed?

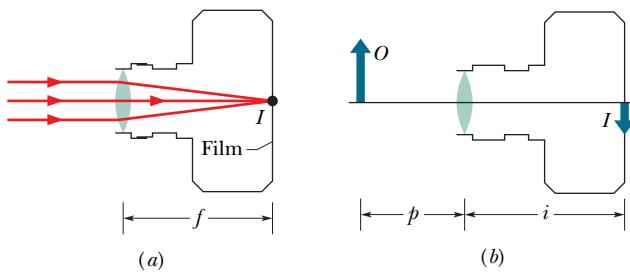


Fig. 34-45 Problem 90.

•91 SSM Figure 34-46a shows the basic structure of a human eye. Light refracts into the eye through the cornea and is then further redirected by a lens whose shape (and thus ability to focus the light) is controlled by muscles. We can treat the cornea and eye lens as a single effective thin lens (Fig. 34-46b). A "normal" eye can focus parallel light rays from a distant object O to a point on the retina at the back of the eye, where processing of the visual information begins. As an object is brought close to the eye, however, the muscles must change the shape of the lens so that rays form an inverted real image on the retina (Fig. 34-46c). (a) Suppose that for the parallel rays of Figs. 34-46a and b, the focal length f of the effective thin lens of the eye is 2.50 cm. For an object at distance $p = 40.0$ cm, what focal length f' of the effective lens is required for the object to be seen clearly? (b)

Must the eye muscles increase or decrease the radii of curvature of the eye lens to give focal length f' ?

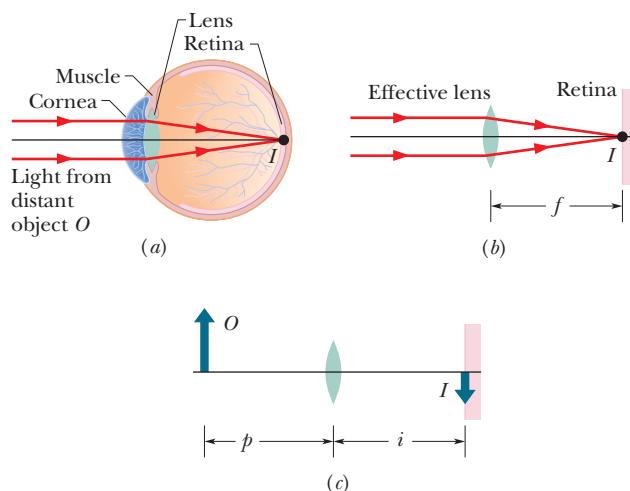


Fig. 34-46 Problem 91.

•92 An object is 10.0 mm from the objective of a certain compound microscope. The lenses are 300 mm apart, and the intermediate image is 50.0 mm from the eyepiece. What overall magnification is produced by the instrument?

•93 Someone with a near point P_n of 25 cm views a thimble through a simple magnifying lens of focal length 10 cm by placing the lens near his eye. What is the angular magnification of the thimble if it is positioned so that its image appears at (a) P_n and (b) infinity?

Additional Problems

94 An object is placed against the center of a spherical mirror and then moved 70 cm from it along the central axis as the image distance i is measured. Figure 34-47 gives i versus object distance p out to $p_s = 40$ cm. What is the image distance when the object is 70 cm from the mirror?

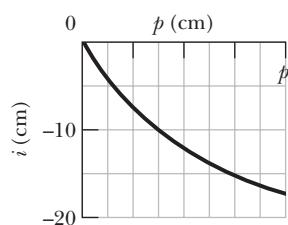


Fig. 34-47 Problem 94.

95 through 100 GO 95, 99 Three-lens systems. In Fig. 34-48, stick figure O (the object) stands on the common central axis of three thin, symmetric lenses, which are mounted in the boxed regions. Lens 1 is mounted within the boxed region closest to O ,

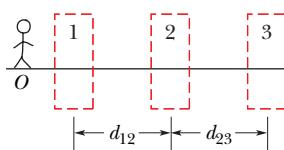


Fig. 34-48 Problems 95 through 100.

Table 34-10

Problems 95 through 100: Three-Lens Systems. See the setup for these problems.

	p_1	Lens 1	d_{12}	Lens 2	d_{23}	Lens 3	(a) i_3	(b) M	(c) R/V	(d) I/NI	(e) Side
95	+12	C, 8.0	28	C, 6.0	8.0	C, 6.0					
96	+4.0	D, 6.0	9.6	C, 6.0	14	C, 4.0					
97	+18	C, 6.0	15	C, 3.0	11	C, 3.0					
98	+2.0	C, 6.0	15	C, 6.0	19	C, 5.0					
99	+8.0	D, 8.0	8.0	D, 16	5.1	C, 8.0					
100	+4.0	C, 6.0	8.0	D, 4.0	5.7	D, 12					

which is at object distance p_1 . Lens 2 is mounted within the middle boxed region, at distance d_{12} from lens 1. Lens 3 is mounted in the farthest boxed region, at distance d_{23} from lens 2. Each problem in Table 34-10 refers to a different combination of lenses and different values for distances, which are given in centimeters. The type of lens is indicated by C for converging and D for diverging; the number after C or D is the distance between a lens and either of the focal points (the proper sign of the focal distance is not indicated).

Find (a) the image distance i_3 for the (final) image produced by lens 3 (the final image produced by the system) and (b) the overall lateral magnification M for the system, including signs. Also, determine whether the final image is (c) real (R) or virtual (V), (d) inverted (I) from object O or noninverted (NI), and (e) on the same side of lens 3 as object O or on the opposite side.

101 SSM The formula $1/p + 1/i = 1/f$ is called the *Gaussian* form of the thin-lens formula. Another form of this formula, the *Newtonian* form, is obtained by considering the distance x from the object to the first focal point and the distance x' from the second focal point to the image. Show that $xx' = f^2$ is the Newtonian form of the thin-lens formula.

102 Figure 34-49a is an overhead view of two vertical plane mirrors with an object O placed between them. If you look into the mirrors, you see multiple images of O . You can find them by drawing the reflection in each mirror of the angular region between the mirrors, as is done in Fig. 34-49b for the left-hand mirror. Then draw the reflection of the reflection. Continue this on the left and on the right until the reflections meet or overlap at the rear of the mirrors. Then you can count the number of images of O . How many images of O would you see if (a) 90° , (b) 45° , and (c) 60° ? If $\theta = 120^\circ$, determine the (d) smallest and (e) largest number of images that can be seen, depending on your perspective and the location of O . (f) In each situation, draw the image locations and orientations as in Fig. 34-49b.

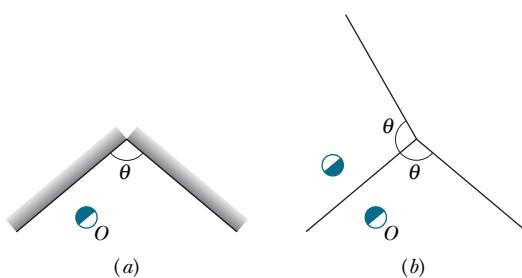


Fig. 34-49 Problem 102.

103 SSM Two thin lenses of focal lengths f_1 and f_2 are in contact. Show that they are equivalent to a single thin lens for which the focal length is $f = f_1 f_2 / (f_1 + f_2)$.

104 Two plane mirrors are placed parallel to each other and 40 cm apart. An object is placed 10 cm from one mirror. Determine the (a) smallest, (b) second smallest, (c) third smallest (occurs twice), and (d) fourth smallest distance between the object and image of the object.

105 In Fig. 34-50, a box is somewhere at the left, on the central axis of the thin converging lens. The image I_m of the box produced by the plane mirror is 4.00 cm “inside” the mirror. The lens-mirror separation is 10.0 cm, and the focal length of the lens is 2.00 cm. (a) What is the distance between the box and the lens? Light reflected by the mirror travels back through the lens, which produces a final image of the box. (b) What is the distance between the lens and that final image?

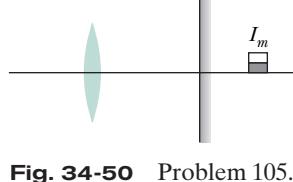


Fig. 34-50 Problem 105.

106 In Fig. 34-51, an object is placed in front of a converging lens at a distance equal to twice the focal length f_1 of the lens. On the other side of the lens is a concave mirror of focal length f_2 separated from the lens by a distance $2(f_1 + f_2)$. Light from the object passes rightward through the lens, reflects from the mirror, passes leftward through the lens, and forms a final image of the object. What are (a) the distance between the lens and that final image and (b) the overall lateral magnification M of the object? Is the image (c) real or virtual (if it is virtual, it requires someone looking through the lens toward the mirror), (d) to the left or right of the lens, and (e) inverted or noninverted relative to the object?

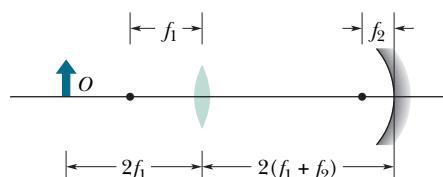


Fig. 34-51 Problem 106.

107 SSM A fruit fly of height H sits in front of lens 1 on the central axis through the lens. The lens forms an image of the fly at a distance $d = 20$ cm from the fly; the image has the fly's orientation and height $H_I = 2.0H$. What are (a) the focal length f_1 of the lens and (b) the object distance p_1 of the fly? The fly then leaves lens

PROBLEMS

957

1 and sits in front of lens 2, which also forms an image at $d = 20\text{ cm}$ that has the same orientation as the fly, but now $H_I = 0.50H$. What are (c) f_2 and (d) p_2 ?

108 You grind the lenses shown in Fig. 34-52 from flat glass disks ($n = 1.5$) using a machine that can grind a radius of curvature of either 40 cm or 60 cm. In a lens where either radius is appropriate, you select the 40 cm radius. Then you hold each lens in sunshine to form an image of the Sun. What are the (a) focal length f and (b) image type (real or virtual) for (bi-convex) lens 1, (c) f and (d) image type for (plane-convex) lens 2, (e) f and (f) image type for (meniscus convex) lens 3, (g) f and (h) image type for (bi-concave) lens 4, (i) f and (j) image type for (plane-concave) lens 5, and (k) f and (l) image type for (meniscus concave) lens 6?

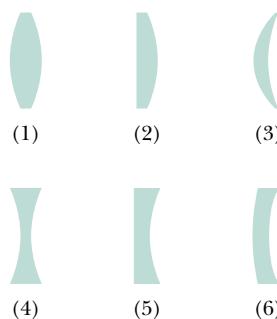


Fig. 34-52
Problem 108.

109 In Fig. 34-53, a fish watcher at point P watches a fish through a glass wall of a fish tank. The watcher is level with the fish; the index of refraction of the glass is $8/5$, and that of the water is $4/3$. The distances are $d_1 = 8.0\text{ cm}$, $d_2 = 3.0\text{ cm}$, and $d_3 = 6.8\text{ cm}$. (a) To the fish, how far away does the watcher appear to be? (Hint: The watcher is the object. Light from that object passes through the wall's outside surface, which acts as a refracting surface. Find the image produced by that surface. Then treat that image as an object whose light passes through the wall's inside surface, which acts as another refracting surface. Find the image produced by that surface, and there is the answer.) (b) To the watcher, how far away does the fish appear to be?

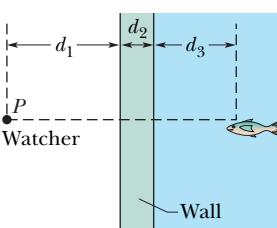


Fig. 34-53
Problem 109.

110 A goldfish in a spherical fish bowl of radius R is at the level of the center C of the bowl and at distance $R/2$ from the glass (Fig. 34-54). What magnification of the fish is produced by the water in

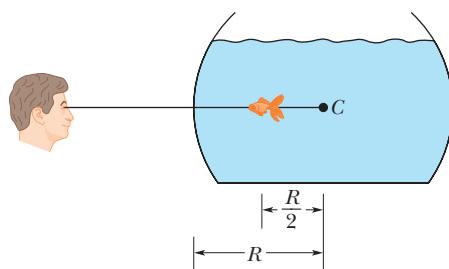


Fig. 34-54 Problem 110.

the bowl for a viewer looking along a line that includes the fish and the center, with the fish on the near side of the center? The index of refraction of the water is 1.33 . Neglect the glass wall of the bowl. Assume the viewer looks with one eye. (Hint: Equation 34-5 holds, but Eq. 34-6 does not. You need to work with a ray diagram of the situation and assume that the rays are close to the observer's line of sight—that is, they deviate from that line by only small angles.)

111 Figure 34-55 shows a *beam expander* made with two coaxial converging lenses of focal lengths f_1 and f_2 and separation $d = f_1 + f_2$. The device can expand a laser beam while keeping the light rays in the beam parallel to the central axis through the lenses. Suppose a uniform laser beam of width $W_i = 2.5\text{ mm}$ and intensity $I_i = 9.0\text{ kW/m}^2$ enters a beam expander for which $f_1 = 12.5\text{ cm}$ and $f_2 = 30.0\text{ cm}$. What are (a) W_f and (b) I_f of the beam leaving the expander? (c) What value of d is needed for the beam expander if lens 1 is replaced with a diverging lens of focal length $f_1 = -26.0\text{ cm}$?

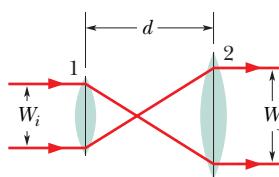


Fig. 34-55 Problem 111.

112 You look down at a coin that lies at the bottom of a pool of liquid of depth d and index of refraction n (Fig. 34-56). Because you view with two eyes, which intercept different rays of light from the coin, you perceive the coin to be where extensions of the intercepted rays cross, at depth d_a instead of d . Assuming that the intercepted rays in Fig. 34-56 are close to a vertical axis through the coin, show that $d_a = d/n$. (Hint: Use the small-angle approximation $\sin \theta \approx \tan \theta \approx \theta$.)

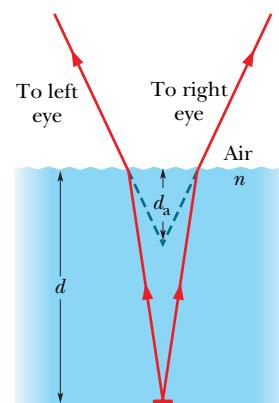


Fig. 34-56 Problem 112.

35

INTERFERENCE

35-1 WHAT IS PHYSICS?

One of the major goals of physics is to understand the nature of light. This goal has been difficult to achieve (and has not yet fully been achieved) because light is complicated. However, this complication means that light offers many opportunities for applications, and some of the richest opportunities involve the interference of light waves—**optical interference**.

Nature has long used optical interference for coloring. For example, the wings of a *Morpho* butterfly are a dull, uninspiring brown, as can be seen on the bottom wing surface, but the brown is hidden on the top surface by an arresting blue due to the interference of light reflecting from that surface (Fig. 35-1). Moreover, the top surface is color-shifting; if you change your perspective or if the wing moves, the tint of the color changes. Similar color shifting is used in the inks on many currencies to thwart counterfeiters, whose copy machines can duplicate color from only one perspective and therefore cannot duplicate any shift in color caused by a change in perspective.

To understand the basic physics of optical interference, we must largely abandon the simplicity of geometrical optics (in which we describe light as rays) and return to the wave nature of light.



Fig. 35-1 The blue of the top surface of a *Morpho* butterfly wing is due to optical interference and shifts in color as your viewing perspective changes. (*Philippe Colombe/PhotoDisc/Getty Images*)

35-2 Light as a Wave

The first person to advance a convincing wave theory for light was Dutch physicist Christian Huygens, in 1678. Although much less comprehensive than the later electromagnetic theory of Maxwell, Huygens' theory was simpler mathematically and remains useful today. Its great advantages are that it accounts for the laws of reflection and refraction in terms of waves and gives physical meaning to the index of refraction.

Huygens' wave theory is based on a geometrical construction that allows us to tell where a given wavefront will be at any time in the future if we know its present position. This construction is based on **Huygens' principle**, which is:

 All points on a wavefront serve as point sources of spherical secondary wavelets. After a time t , the new position of the wavefront will be that of a surface tangent to these secondary wavelets.

Here is a simple example. At the left in Fig. 35-2, the present location of a wavefront of a plane wave traveling to the right in vacuum is represented by plane ab , perpendicular to the page. Where will the wavefront be at time Δt later? We let several points on plane ab (the dots) serve as sources of spherical secondary wavelets that are emitted at $t = 0$. At time Δt , the radius of all these spherical wavelets will have grown to $c \Delta t$, where c is the speed of light in vacuum. We draw plane de tangent to these wavelets at time Δt . This plane represents the wavefront of the plane wave at time Δt ; it is parallel to plane ab and a perpendicular distance $c \Delta t$ from it.

The Law of Refraction

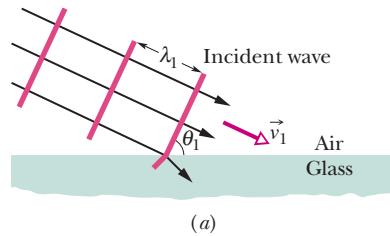
We now use Huygens' principle to derive the law of refraction, Eq. 33-40 (Snell's law). Figure 35-3 shows three stages in the refraction of several wavefronts at a flat interface between air (medium 1) and glass (medium 2). We arbitrarily choose the wavefronts in the incident light beam to be separated by λ_1 , the wavelength in medium 1. Let the speed of light in air be v_1 and that in glass be v_2 . We assume that $v_2 < v_1$, which happens to be true.

Angle θ_1 in Fig. 35-3a is the angle between the wavefront and the interface; it has the same value as the angle between the *normal* to the wavefront (that is, the incident ray) and the *normal* to the interface. Thus, θ_1 is the angle of incidence.

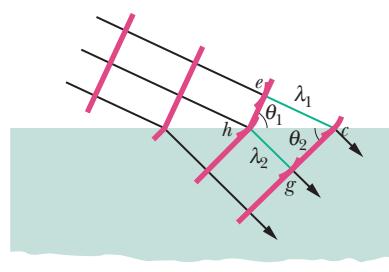
As the wave moves into the glass, a Huygens wavelet at point e in Fig. 35-3b will expand to pass through point c , at a distance of λ_1 from point e . The time interval required for this expansion is that distance divided by the speed of the wavelet, or λ_1/v_1 . Now note that in this same time interval, a Huygens wavelet at point h will expand to pass through point g , at the reduced speed v_2 and with wavelength λ_2 . Thus, this time interval must also be equal to λ_2/v_2 . By equating these times of travel, we obtain the relation

$$\frac{\lambda_1}{\lambda_2} = \frac{v_1}{v_2}, \quad (35-1)$$

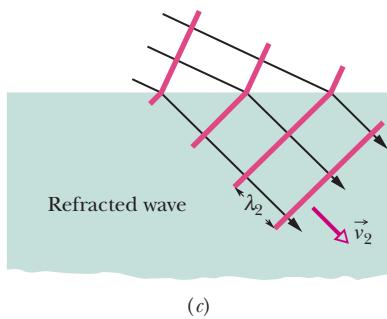
Refraction occurs at the surface, giving a new direction of travel.



(a)



(b)



(c)

Fig. 35-3 The refraction of a plane wave at an air–glass interface, as portrayed by Huygens' principle. The wavelength in glass is smaller than that in air. For simplicity, the reflected wave is not shown. Parts (a) through (c) represent three successive stages of the refraction.

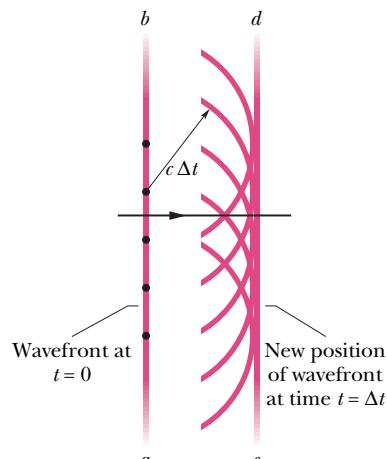


Fig. 35-2 The propagation of a plane wave in vacuum, as portrayed by Huygens' principle.

which shows that the wavelengths of light in two media are proportional to the speeds of light in those media.

By Huygens' principle, the refracted wavefront must be tangent to an arc of radius λ_2 centered on h , say at point g . The refracted wavefront must also be tangent to an arc of radius λ_1 centered on e , say at c . Then the refracted wavefront must be oriented as shown. Note that θ_2 , the angle between the refracted wavefront and the interface, is actually the angle of refraction.

For the right triangles hce and hcg in Fig. 35-3b we may write

$$\sin \theta_1 = \frac{\lambda_1}{hc} \quad (\text{for triangle } hce)$$

and $\sin \theta_2 = \frac{\lambda_2}{hc} \quad (\text{for triangle } hcg).$

Dividing the first of these two equations by the second and using Eq. 35-1, we find

$$\frac{\sin \theta_1}{\sin \theta_2} = \frac{\lambda_1}{\lambda_2} = \frac{v_1}{v_2}. \quad (35-2)$$

We can define the **index of refraction** n for each medium as the ratio of the speed of light in vacuum to the speed of light v in the medium. Thus,

$$n = \frac{c}{v} \quad (\text{index of refraction}). \quad (35-3)$$

In particular, for our two media, we have

$$n_1 = \frac{c}{v_1} \quad \text{and} \quad n_2 = \frac{c}{v_2}. \quad (35-4)$$

If we combine Eqs. 35-2 and 35-4, we find

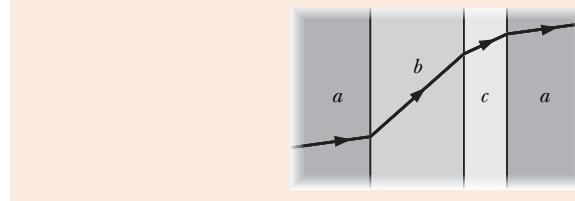
$$\frac{\sin \theta_1}{\sin \theta_2} = \frac{c/n_1}{c/n_2} = \frac{n_2}{n_1} \quad (35-5)$$

or $n_1 \sin \theta_1 = n_2 \sin \theta_2 \quad (\text{law of refraction}), \quad (35-6)$

as introduced in Chapter 33.

CHECKPOINT 1

The figure shows a monochromatic ray of light traveling across parallel interfaces, from an original material a , through layers of materials b and c , and then back into material a . Rank the materials according to the speed of light in them, greatest first.



Wavelength and Index of Refraction

We have now seen that the wavelength of light changes when the speed of the light changes, as happens when light crosses an interface from one medium into another. Further, the speed of light in any medium depends on the index of refraction of the medium, according to Eq. 35-3. Thus, the wavelength of light in any medium depends on the index of refraction of the medium. Let a certain

monochromatic light have wavelength λ and speed c in vacuum and wavelength λ_n and speed v in a medium with an index of refraction n . Now we can rewrite Eq. 35-1 as

$$\lambda_n = \lambda \frac{v}{c}. \quad (35-7)$$

Using Eq. 35-3 to substitute $1/n$ for v/c then yields

$$\lambda_n = \frac{\lambda}{n}. \quad (35-8)$$

This equation relates the wavelength of light in any medium to its wavelength in vacuum. It tells us that the greater the index of refraction of a medium, the smaller the wavelength of light in that medium.

What about the frequency of the light? Let f_n represent the frequency of the light in a medium with index of refraction n . Then from the general relation of Eq. 16-13 ($v = \lambda f$), we can write

$$f_n = \frac{v}{\lambda_n}.$$

Substituting Eqs. 35-3 and 35-8 then gives us

$$f_n = \frac{c/n}{\lambda/n} = \frac{c}{\lambda} = f,$$

where f is the frequency of the light in vacuum. Thus, although the speed and wavelength of light in the medium are different from what they are in vacuum, *the frequency of the light in the medium is the same as it is in vacuum*.

The fact that the wavelength of light depends on the index of refraction via Eq. 35-8 is important in certain situations involving the interference of light waves. For example, in Fig. 35-4, the *waves of the rays* (that is, the waves represented by the rays) have identical wavelengths λ and are initially in phase in air ($n \approx 1$). One of the waves travels through medium 1 of index of refraction n_1 and length L . The other travels through medium 2 of index of refraction n_2 and the same length L . When the waves leave the two media, they will have the same wavelength—their wavelength λ in air. However, because their wavelengths differed in the two media, the two waves may no longer be in phase.



The phase difference between two light waves can change if the waves travel through different materials having different indexes of refraction.

As we shall discuss soon, this change in the phase difference can determine how the light waves will interfere if they reach some common point.

To find their new phase difference in terms of wavelengths, we first count the number N_1 of wavelengths there are in the length L of medium 1. From Eq. 35-8, the wavelength in medium 1 is $\lambda_{n1} = \lambda/n_1$; so

$$N_1 = \frac{L}{\lambda_{n1}} = \frac{Ln_1}{\lambda}. \quad (35-9)$$

Similarly, we count the number N_2 of wavelengths there are in the length L of medium 2, where the wavelength is $\lambda_{n2} = \lambda/n_2$:

$$N_2 = \frac{L}{\lambda_{n2}} = \frac{Ln_2}{\lambda}. \quad (35-10)$$

To find the new phase difference between the waves, we subtract the smaller of N_1 and N_2 from the larger. Assuming $n_2 > n_1$, we obtain

$$N_2 - N_1 = \frac{Ln_2}{\lambda} - \frac{Ln_1}{\lambda} = \frac{L}{\lambda} (n_2 - n_1). \quad (35-11)$$

The difference in indexes causes a phase shift between the rays.

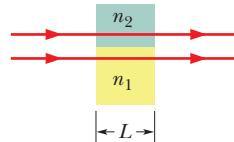


Fig. 35-4 Two light rays travel through two media having different indexes of refraction.

Suppose Eq. 35-11 tells us that the waves now have a phase difference of 45.6 wavelengths. That is equivalent to taking the initially in-phase waves and shifting one of them by 45.6 wavelengths. However, a shift of an integer number of wavelengths (such as 45) would put the waves back in phase; so it is only the decimal fraction (here, 0.6) that is important. A phase difference of 45.6 wavelengths is equivalent to an *effective phase difference* of 0.6 wavelength.

A phase difference of 0.5 wavelength puts two waves exactly out of phase. If the waves had equal amplitudes and were to reach some common point, they would then undergo fully destructive interference, producing darkness at that point. With a phase difference of 0.0 or 1.0 wavelength, they would, instead, undergo fully constructive interference, resulting in brightness at the common point. Our phase difference of 0.6 wavelength is an intermediate situation but closer to fully destructive interference, and the waves would produce a dimly illuminated common point.

We can also express phase difference in terms of radians and degrees, as we have done already. A phase difference of one wavelength is equivalent to phase differences of 2π rad and 360° .

CHECKPOINT 2

The light waves of the rays in Fig. 35-4 have the same wavelength and amplitude and are initially in phase. (a) If 7.60 wavelengths fit within the length of the top material and 5.50 wavelengths fit within that of the bottom material, which material has the greater index of refraction? (b) If the rays are angled slightly so that they meet at the same point on a distant screen, will the interference there result in the brightest possible illumination, bright intermediate illumination, dark intermediate illumination, or darkness?

Rainbows and Optical Interference

In Section 33-8, we discussed how the colors of sunlight are separated into a rainbow when sunlight travels through falling raindrops. We dealt with a simplified situation in which a single ray of white light entered a drop. Actually, light waves pass into a drop along the entire side that faces the Sun. Here we cannot discuss the details of how these waves travel through the drop and then emerge, but we can see that different parts of an incoming wave will travel different paths within the drop. That means waves will emerge from the drop with different phases. Thus, we can see that at some angles the emerging light will be in phase and give constructive interference. The rainbow is the result of such constructive interference. For example, the red of the rainbow appears because waves of red light emerge in phase from each raindrop in the direction in which you see that part of the rainbow. The light waves that emerge in other directions from each raindrop have a range of different phases because they take a range of different paths through each drop. This light is neither bright nor colorful, and so you do not notice it.

If you are lucky and look carefully below a primary rainbow, you can see dimmer colored arcs called *supernumeraries* (Fig. 35-5). Like the main arcs of the rainbow, the supernumeraries are due to waves that emerge from each drop approximately in phase with one another to give constructive interference. If you are very lucky and look very carefully above a secondary rainbow, you might see even more (but even dimmer) supernumeraries. Keep in mind that both types of rainbows and both sets of supernumeraries are naturally occurring examples of optical interference and naturally occurring evidence that light consists of waves.

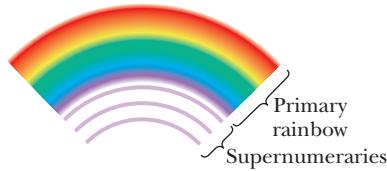


Fig. 35-5 A primary rainbow and the faint supernumeraries below it are due to optical interference.

Sample Problem**Phase difference of two waves due to difference in refractive indexes**

In Fig. 35-4, the two light waves that are represented by the rays have wavelength 550.0 nm before entering media 1 and 2. They also have equal amplitudes and are in phase. Medium 1 is now just air, and medium 2 is a transparent plastic layer of index of refraction 1.600 and thickness 2.600 μm .

- (a) What is the phase difference of the emerging waves in wavelengths, radians, and degrees? What is their effective phase difference (in wavelengths)?

KEY IDEA

The phase difference of two light waves can change if they travel through different media, with different indexes of refraction. The reason is that their wavelengths are different in the different media. We can calculate the change in phase difference by counting the number of wavelengths that fits into each medium and then subtracting those numbers.

Calculations: When the path lengths of the waves in the two media are identical, Eq. 35-11 gives the result of the subtraction. Here we have $n_1 = 1.000$ (for the air), $n_2 = 1.600$, $L = 2.600 \mu\text{m}$, and $\lambda = 550.0 \text{ nm}$. Thus, Eq. 35-11 yields

$$\begin{aligned} N_2 - N_1 &= \frac{L}{\lambda} (n_2 - n_1) \\ &= \frac{2.600 \times 10^{-6} \text{ m}}{5.500 \times 10^{-7} \text{ m}} (1.600 - 1.000) \\ &= 2.84. \end{aligned} \quad (\text{Answer})$$



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Thus, the phase difference of the emerging waves is 2.84 wavelengths. Because 1.0 wavelength is equivalent to $2\pi \text{ rad}$ and 360° , you can show that this phase difference is equivalent to

$$\text{phase difference} = 17.8 \text{ rad} \approx 1020^\circ. \quad (\text{Answer})$$

The effective phase difference is the decimal part of the actual phase difference *expressed in wavelengths*. Thus, we have

$$\text{effective phase difference} = 0.84 \text{ wavelength}. \quad (\text{Answer})$$

You can show that this is equivalent to 5.3 rad and about 300° . *Caution:* We do *not* find the effective phase difference by taking the decimal part of the actual phase difference as expressed in radians or degrees. For example, we do *not* take 0.8 rad from the actual phase difference of 17.8 rad.

- (b) If the waves reached the same point on a distant screen, what type of interference would they produce?

Reasoning: We need to compare the effective phase difference of the waves with the phase differences that give the extreme types of interference. Here the effective phase difference of 0.84 wavelength is between 0.5 wavelength (for fully destructive interference, or the darkest possible result) and 1.0 wavelength (for fully constructive interference, or the brightest possible result), but closer to 1.0 wavelength. Thus, the waves would produce intermediate interference that is closer to fully constructive interference—they would produce a relatively bright spot.

35-3 Diffraction

In the next section we shall discuss the experiment that first proved that light is a wave. To prepare for that discussion, we must introduce the idea of **diffraction** of waves, a phenomenon that we explore much more fully in Chapter 36. Its essence is this: If a wave encounters a barrier that has an opening of dimensions similar to the wavelength, the part of the wave that passes through the opening will flare (spread) out—will *diffract*—into the region beyond the barrier. The flaring is consistent with the spreading of wavelets in the Huygens construction of Fig. 35-2. Diffraction occurs for waves of all types, not just light waves; Fig. 35-6 shows the diffraction of water waves traveling across the surface of water in a shallow tank.



Image restricted for electronic distribution

Fig. 35-6 Waves produced by an oscillating paddle at the left flare out through an opening in a barrier along the water surface. (Runk Schoenberger/Grant Heilman Photography)

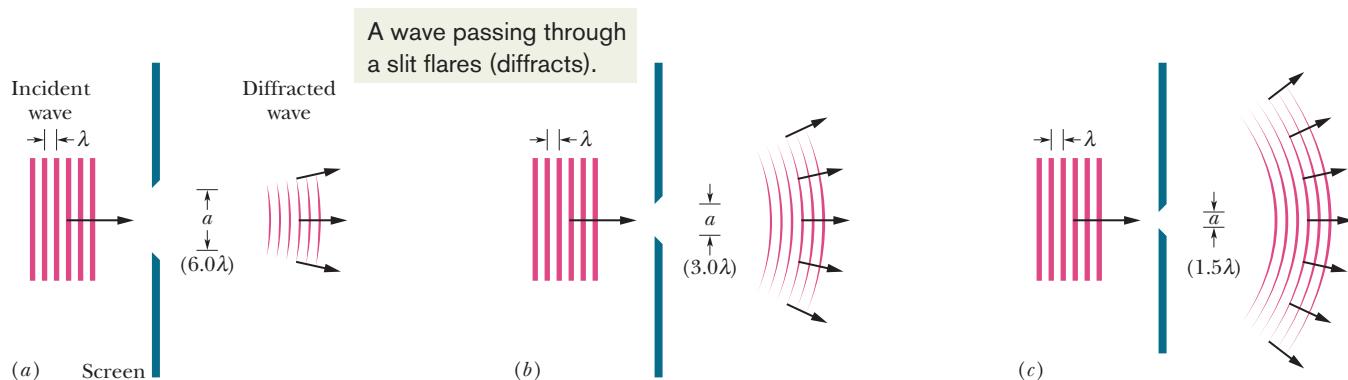


Fig. 35-7 Diffraction represented schematically. For a given wavelength λ , the diffraction is more pronounced the smaller the slit width a . The figures show the cases for (a) slit width $a = 6.0\lambda$, (b) slit width $a = 3.0\lambda$, and (c) slit width $a = 1.5\lambda$. In all three cases, the screen and the length of the slit extend well into and out of the page, perpendicular to it.

Figure 35-7a shows the situation schematically for an incident plane wave of wavelength λ encountering a slit that has width $a = 6.0\lambda$ and extends into and out of the page. The part of the wave that passes through the slit flares out on the far side. Figures 35-7b (with $a = 3.0\lambda$) and 35-7c ($a = 1.5\lambda$) illustrate the main feature of diffraction: the narrower the slit, the greater the diffraction.

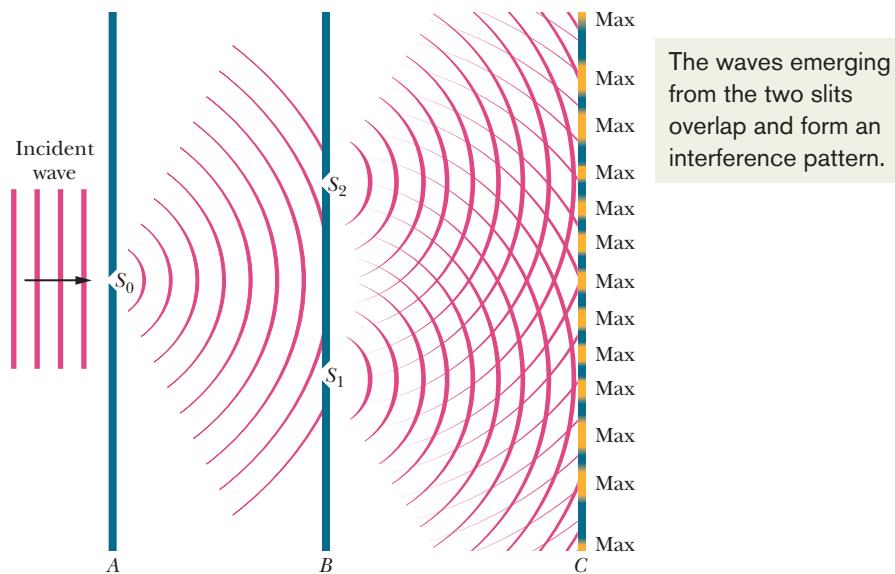
Diffraction limits geometrical optics, in which we represent an electromagnetic wave with a ray. If we actually try to form a ray by sending light through a narrow slit, or through a series of narrow slits, diffraction will always defeat our effort because it always causes the light to spread. Indeed, the narrower we make the slits (in the hope of producing a narrower beam), the greater the spreading is. Thus, geometrical optics holds only when slits or other apertures that might be located in the path of light do not have dimensions comparable to or smaller than the wavelength of the light.

35-4 Young's Interference Experiment

In 1801, Thomas Young experimentally proved that light is a wave, contrary to what most other scientists then thought. He did so by demonstrating that light undergoes interference, as do water waves, sound waves, and waves of all other types. In addition, he was able to measure the average wavelength of sunlight; his value, 570 nm, is impressively close to the modern accepted value of 555 nm. We shall here examine Young's experiment as an example of the interference of light waves.

Figure 35-8 gives the basic arrangement of Young's experiment. Light from a distant monochromatic source illuminates slit S_0 in screen A. The emerging light

Fig. 35-8 In Young's interference experiment, incident monochromatic light is diffracted by slit S_0 , which then acts as a point source of light that emits semicircular wavefronts. As that light reaches screen B , it is diffracted by slits S_1 and S_2 , which then act as two point sources of light. The light waves traveling from slits S_1 and S_2 overlap and undergo interference, forming an interference pattern of maxima and minima on viewing screen C . This figure is a cross section; the screens, slits, and interference pattern extend into and out of the page. Between screens B and C , the semicircular wavefronts centered on S_2 depict the waves that would be there if only S_2 were open. Similarly, those centered on S_1 depict waves that would be there if only S_1 were open.



then spreads via diffraction to illuminate two slits S_1 and S_2 in screen B . Diffraction of the light by these two slits sends overlapping circular waves into the region beyond screen B , where the waves from one slit interfere with the waves from the other slit.

The “snapshot” of Fig. 35-8 depicts the interference of the overlapping waves. However, we cannot see evidence for the interference except where a viewing screen C intercepts the light. Where it does so, points of interference maxima form visible bright rows—called *bright bands*, *bright fringes*, or (loosely speaking) *maxima*—that extend across the screen (into and out of the page in Fig. 35-8). Dark regions—called *dark bands*, *dark fringes*, or (loosely speaking) *minima*—result from fully destructive interference and are visible between adjacent pairs of bright fringes. (*Maxima* and *minima* more properly refer to the center of a band.) The pattern of bright and dark fringes on the screen is called an **interference pattern**. Figure 35-9 is a photograph of part of the interference pattern that would be seen by an observer standing to the left of screen C in the arrangement of Fig. 35-8.

Locating the Fringes

Light waves produce fringes in a *Young's double-slit interference experiment*, as it is called, but what exactly determines the locations of the fringes? To answer, we shall use the arrangement in Fig. 35-10a. There, a plane wave of monochromatic light is incident on two slits S_1 and S_2 in screen B ; the light diffracts through the slits and produces an interference pattern on screen C . We draw a central axis from the point halfway between the slits to screen C as a reference. We then pick, for discussion, an arbitrary point P on the screen, at angle θ to the central axis. This point intercepts the wave of ray r_1 from the bottom slit and the wave of ray r_2 from the top slit.

These waves are in phase when they pass through the two slits because there they are just portions of the same incident wave. However, once they have passed the slits, the two waves must travel different distances to reach P . We saw a similar situation in Section 17-5 with sound waves and concluded that



The phase difference between two waves can change if the waves travel paths of different lengths.

The change in phase difference is due to the *path length difference* ΔL in the paths taken by the waves. Consider two waves initially exactly in phase, traveling along paths with a path length difference ΔL , and then passing through some common point. When ΔL is zero or an integer number of wavelengths, the waves arrive at the common point exactly in phase and they interfere fully con-

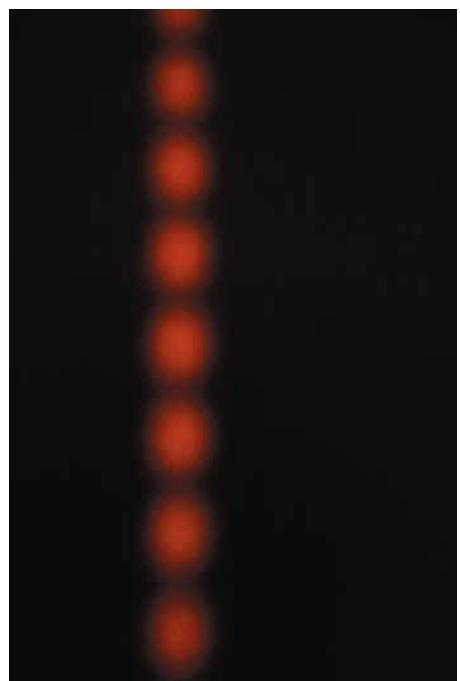


Fig. 35-9 A photograph of the interference pattern produced by the arrangement shown in Fig. 35-8, but with short slits. (The photograph is a front view of part of screen C .) The alternating maxima and minima are called *interference fringes* (because they resemble the decorative fringe sometimes used on clothing and rugs). (Jearl Walker)

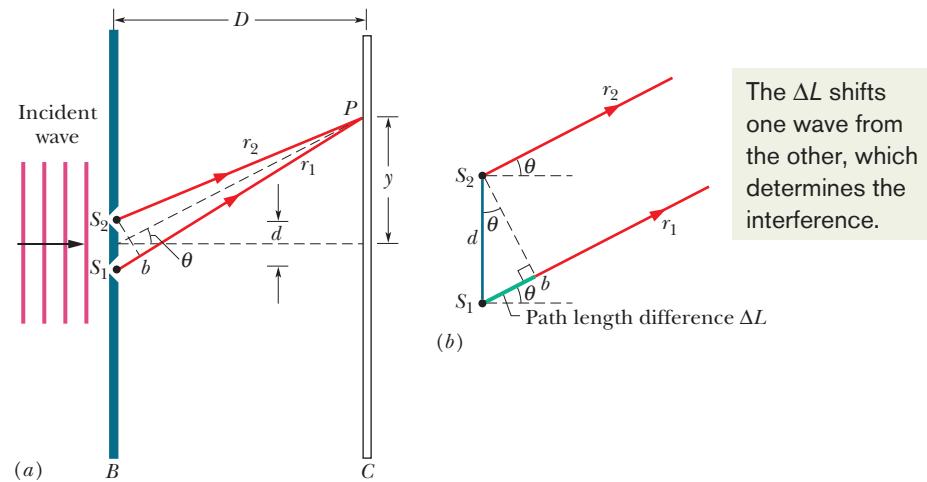


Fig. 35-10 (a) Waves from slits S_1 and S_2 (which extend into and out of the page) combine at P , an arbitrary point on screen C at distance y from the central axis. The angle θ serves as a convenient locator for P . (b) For $D \gg d$, we can approximate rays r_1 and r_2 as being parallel, at angle θ to the central axis.

structively there. If that is true for the waves of rays r_1 and r_2 in Fig. 35-10, then point P is part of a bright fringe. When, instead, ΔL is an odd multiple of half a wavelength, the waves arrive at the common point exactly out of phase and they interfere fully destructively there. If that is true for the waves of rays r_1 and r_2 , then point P is part of a dark fringe. (And, of course, we can have intermediate situations of interference and thus intermediate illumination at P .) Thus,



What appears at each point on the viewing screen in a Young's double-slit interference experiment is determined by the path length difference ΔL of the rays reaching that point.

We can specify where each bright fringe and each dark fringe is located on the screen by giving the angle θ from the central axis to that fringe. To find θ , we must relate it to ΔL . We start with Fig. 35-10a by finding a point b along ray r_1 such that the path length from b to P equals the path length from S_2 to P . Then the path length difference ΔL between the two rays is the distance from S_1 to b .

The relation between this S_1 -to- b distance and θ is complicated, but we can simplify it considerably if we arrange for the distance D from the slits to the screen to be much greater than the slit separation d . Then we can approximate rays r_1 and r_2 as being parallel to each other and at angle θ to the central axis (Fig. 35-10b). We can also approximate the triangle formed by S_1 , S_2 , and b as being a right triangle, and approximate the angle inside that triangle at S_2 as being θ . Then, for that triangle, $\sin \theta = \Delta L/d$ and thus

$$\Delta L = d \sin \theta \quad (\text{path length difference}). \quad (35-12)$$

For a bright fringe, we saw that ΔL must be either zero or an integer number of wavelengths. Using Eq. 35-12, we can write this requirement as

$$\Delta L = d \sin \theta = (\text{integer})(\lambda), \quad (35-13)$$

or as

$$d \sin \theta = m\lambda, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{maxima—bright fringes}). \quad (35-14)$$

For a dark fringe, ΔL must be an odd multiple of half a wavelength. Again using Eq. 35-12, we can write this requirement as

$$\Delta L = d \sin \theta = (\text{odd number})\left(\frac{1}{2}\lambda\right), \quad (35-15)$$

or as

$$d \sin \theta = (m + \frac{1}{2})\lambda, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{minima—dark fringes}). \quad (35-16)$$

With Eqs. 35-14 and 35-16, we can find the angle θ to any fringe and thus locate that fringe; further, we can use the values of m to label the fringes. For the value and label $m = 0$, Eq. 35-14 tells us that a bright fringe is at $\theta = 0$ and thus on the central axis. This *central maximum* is the point at which waves arriving from the two slits have a path length difference $\Delta L = 0$, hence zero phase difference.

For, say, $m = 2$, Eq. 35-14 tells us that *bright* fringes are at the angle

$$\theta = \sin^{-1}\left(\frac{2\lambda}{d}\right)$$

above and below the central axis. Waves from the two slits arrive at these two fringes with $\Delta L = 2\lambda$ and with a phase difference of two wavelengths. These fringes are said to be the *second-order bright fringes* (meaning $m = 2$) or the *second side maxima* (the second maxima to the side of the central maximum), or they are described as being the second bright fringes from the central maximum.

For $m = 1$, Eq. 35-16 tells us that *dark fringes* are at the angle

$$\theta = \sin^{-1}\left(\frac{1.5\lambda}{d}\right)$$

above and below the central axis. Waves from the two slits arrive at these two fringes with $\Delta L = 1.5\lambda$ and with a phase difference, in wavelengths, of 1.5. These fringes are called the *second-order dark fringes* or *second minima* because they are the second dark fringes to the side of the central axis. (The first dark fringes, or first minima, are at locations for which $m = 0$ in Eq. 35-16.)

We derived Eqs. 35-14 and 35-16 for the situation $D \gg d$. However, they also apply if we place a converging lens between the slits and the viewing screen and then move the viewing screen closer to the slits, to the focal point of the lens. (The screen is then said to be in the *focal plane* of the lens; that is, it is in the plane perpendicular to the central axis at the focal point.) One property of a converging lens is that it focuses all rays that are parallel to one another to the same point on its focal plane. Thus, the rays that now arrive at any point on the screen (in the focal plane) were exactly parallel (rather than approximately) when they left the slits. They are like the initially parallel rays in Fig. 34-14a that are directed to a point (the focal point) by a lens.



CHECKPOINT 3

In Fig. 35-10a, what are ΔL (as a multiple of the wavelength) and the phase difference (in wavelengths) for the two rays if point P is (a) a third side maximum and (b) a third minimum?

Sample Problem

Double-slit interference pattern

What is the distance on screen C in Fig. 35-10a between adjacent maxima near the center of the interference pattern? The wavelength λ of the light is 546 nm, the slit separation d is 0.12 mm, and the slit–screen separation D is 55 cm. Assume that θ in Fig. 35-10 is small enough to permit use of the approximations $\sin \theta \approx \tan \theta \approx \theta$, in which θ is expressed in radian measure.

KEY IDEAS

- (1) First, let us pick a maximum with a low value of m to ensure that it is near the center of the pattern. Then, from the geometry of Fig. 35-10a, the maximum's vertical distance y_m from the center of the pattern is related to its angle θ from the central axis by

$$\tan \theta \approx \theta = \frac{y_m}{D}.$$

- (2) From Eq. 35-14, this angle θ for the m th maximum is given by

$$\sin \theta \approx \theta = \frac{m\lambda}{d}.$$

Calculations: If we equate our two expressions for angle θ and then solve for y_m , we find

$$y_m = \frac{m\lambda D}{d}. \quad (35-17)$$

For the next maximum as we move away from the pattern's center, we have

$$y_{m+1} = \frac{(m+1)\lambda D}{d}. \quad (35-18)$$

We find the distance between these adjacent maxima by subtracting Eq. 35-17 from Eq. 35-18:

$$\begin{aligned} \Delta y &= y_{m+1} - y_m = \frac{\lambda D}{d} \\ &= \frac{(546 \times 10^{-9} \text{ m})(55 \times 10^{-2} \text{ m})}{0.12 \times 10^{-3} \text{ m}} \\ &= 2.50 \times 10^{-3} \text{ m} \approx 2.5 \text{ mm}. \end{aligned} \quad (\text{Answer})$$

As long as d and θ in Fig. 35-10a are small, the separation of the interference fringes is independent of m ; that is, the fringes are evenly spaced.



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Sample Problem

Double-slit interference pattern with plastic over one slit

A double-slit interference pattern is produced on a screen, as in Fig. 35-10; the light is monochromatic at a wavelength of 600 nm. A strip of transparent plastic with index of refraction $n = 1.50$ is to be placed over one of the slits. Its presence changes the interference between light waves from the two slits, causing the interference pattern to be shifted across the screen from the original pattern. Figure 35-11a shows the original locations of the central bright fringe ($m = 0$) and the first bright fringes ($m = 1$) above and below the central fringe. The purpose of the plastic is to shift the pattern upward so that the lower $m = 1$ bright fringe is shifted to the center of the pattern. Should the plastic be placed over the top slit (as arbitrarily drawn in Fig. 35-11b) or the bottom slit, and what thickness L should it have?

KEY IDEA

The interference at a point on the screen depends on the phase difference of the light rays arriving from the two slits. The light rays are in phase at the slits, but their relative phase can shift on the way to the screen due to (1) a difference in the length of the paths they follow and (2) a difference in the number of their internal wavelengths λ_n in the materials through which they pass. The first condition applies to any off-center point, and the second condition applies when the plastic covers one of the slits.

Path length difference: Figure 35-11a shows rays r_1 and r_2 along which waves from the two slits travel to reach the lower $m = 1$ bright fringe. Those waves start in phase at the slits but arrive at the fringe with a phase difference of exactly 1 wavelength. To remind ourselves of this main characteristic of the fringe, let us call it the 1λ fringe. The one-wavelength phase difference is due to the one-wavelength path length difference between the rays reaching the fringe; that is, there is exactly one more wavelength along ray r_2 than along r_1 .

Figure 35-11b shows the 1λ fringe shifted up to the center of the pattern with the plastic strip over the top slit (we still do not know whether the plastic should be there or over the bottom slit). The figure also shows the new orientations of rays r_1 and r_2 to reach that fringe. There still must be one more wavelength along r_2 than along r_1 (because they still produce the 1λ fringe), but now the path length difference between those rays is zero, as we can tell from the geometry of Fig. 35-11b. However, r_2 now passes through the plastic.

Internal wavelength: The wavelength λ_n of light in a material with index of refraction n is smaller than the wavelength in vacuum, as given by Eq. 35-8 ($\lambda_n = \lambda/n$). Here, this means that the wavelength of the light is smaller in the plastic than in the air. Thus, the ray that passes through the plastic will have more wavelengths along it than the ray that passes through only air—so we do get the one extra wavelength we need along ray r_2 by placing the plastic over the top slit, as drawn in Fig. 35-11b.

Thickness: To determine the required thickness L of the plastic, we first note that the waves are initially in phase and travel equal distances L through different materials (plastic and air). Because we know the phase difference and require L , we use Eq. 35-11,

$$N_2 - N_1 = \frac{L}{\lambda} (n_2 - n_1). \quad (35-19)$$

We know that $N_2 - N_1$ is 1 for a phase difference of one wavelength, n_2 is 1.50 for the plastic in front of the top slit, n_1 is 1.00 for the air in front of the bottom slit, and λ is 600×10^{-9} m. Then Eq. 35-19 tells us that, to shift the lower $m = 1$ bright fringe up to the center of the interference pattern, the plastic must have the thickness

$$L = \frac{\lambda(N_2 - N_1)}{n_2 - n_1} = \frac{(600 \times 10^{-9} \text{ m})(1)}{1.50 - 1.00} = 1.2 \times 10^{-6} \text{ m}. \quad (\text{Answer})$$

The difference in indexes causes a phase shift between the rays, moving the 1λ fringe upward.

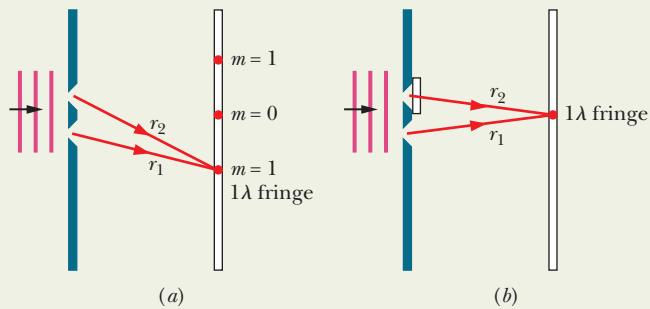


Fig. 35-11 (a) Arrangement for two-slit interference (not to scale). The locations of three bright fringes (or maxima) are indicated. (b) A strip of plastic covers the top slit. We want the 1λ fringe to be at the center of the pattern.

35-5 Coherence

For the interference pattern to appear on viewing screen C in Fig. 35-8, the light waves reaching any point P on the screen must have a phase difference that does not vary in time. That is the case in Fig. 35-8 because the waves passing through slits S_1 and S_2 are portions of the single light wave that illuminates the slits. Because the phase difference remains constant, the light from slits S_1 and S_2 is said to be completely **coherent**.

Direct sunlight is partially coherent; that is, sunlight waves intercepted at two points have a constant phase difference only if the points are very close. If you look closely at your fingernail in bright sunlight, you can see a faint interference pattern called *speckle* that causes the nail to appear to be covered with specks. You see this effect because light waves scattering from very close points on the nail are sufficiently coherent to interfere with one another at your eye. The slits in a double-slit experiment, however, are not close enough, and in direct sunlight, the light at the slits would be **incoherent**. To get coherent light, we would have to send the sunlight through a single slit as in Fig. 35-8; because that single slit is small, light that passes through it is coherent. In addition, the smallness of the slit causes the coherent light to spread via diffraction to illuminate both slits in the double-slit experiment.

If we replace the double slits with two similar but independent monochromatic light sources, such as two fine incandescent wires, the phase difference between the waves emitted by the sources varies rapidly and randomly. (This occurs because the light is emitted by vast numbers of atoms in the wires, acting randomly and independently for extremely short times—of the order of nanoseconds.) As a result, at any given point on the viewing screen, the interference between the waves from the two sources varies rapidly and randomly between fully constructive and fully destructive. The eye (and most common optical detectors) cannot follow such changes, and no interference pattern can be seen. The fringes disappear, and the screen is seen as being uniformly illuminated.

A *laser* differs from common light sources in that its atoms emit light in a cooperative manner, thereby making the light coherent. Moreover, the light is almost monochromatic, is emitted in a thin beam with little spreading, and can be focused to a width that almost matches the wavelength of the light.

35-6 Intensity in Double-Slit Interference

Equations 35-14 and 35-16 tell us how to locate the maxima and minima of the double-slit interference pattern on screen C of Fig. 35-10 as a function of the angle θ in that figure. Here we wish to derive an expression for the intensity I of the fringes as a function of θ .

The light leaving the slits is in phase. However, let us assume that the light waves from the two slits are not in phase when they arrive at point P . Instead, the electric field components of those waves at point P are not in phase and vary with time as

$$E_1 = E_0 \sin \omega t \quad (35-20)$$

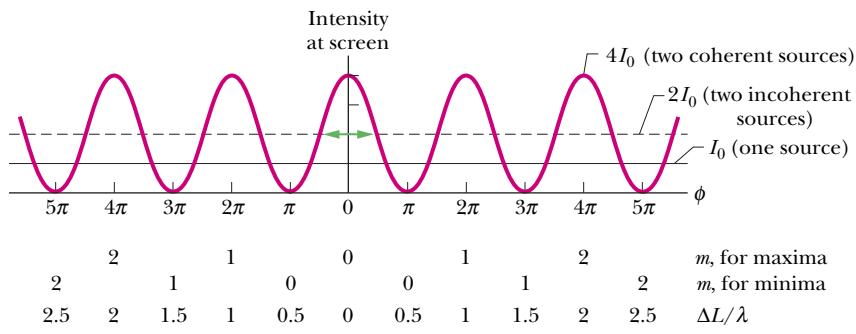
and

$$E_2 = E_0 \sin(\omega t + \phi), \quad (35-21)$$

where ω is the angular frequency of the waves and ϕ is the phase constant of wave E_2 . Note that the two waves have the same amplitude E_0 and a phase difference of ϕ . Because that phase difference does not vary, the waves are coherent. We shall show that these two waves will combine at P to produce an intensity I given by

$$I = 4I_0 \cos^2 \frac{1}{2}\phi, \quad (35-22)$$

Fig. 35-12 A plot of Eq. 35-22, showing the intensity of a double-slit interference pattern as a function of the phase difference between the waves when they arrive from the two slits. I_0 is the (uniform) intensity that would appear on the screen if one slit were covered. The average intensity of the fringe pattern is $2I_0$, and the *maximum* intensity (for coherent light) is $4I_0$.



and that

$$\phi = \frac{2\pi d}{\lambda} \sin \theta. \quad (35-23)$$

In Eq. 35-22, I_0 is the intensity of the light that arrives on the screen from one slit when the other slit is temporarily covered. We assume that the slits are so narrow in comparison to the wavelength that this single-slit intensity is essentially uniform over the region of the screen in which we wish to examine the fringes.

Equations 35-22 and 35-23, which together tell us how the intensity I of the fringe pattern varies with the angle θ in Fig. 35-10, necessarily contain information about the location of the maxima and minima. Let us see if we can extract that information to find equations about those locations.

Study of Eq. 35-22 shows that intensity maxima will occur when

$$\frac{1}{2}\phi = m\pi, \quad \text{for } m = 0, 1, 2, \dots \quad (35-24)$$

If we put this result into Eq. 35-23, we find

$$2m\pi = \frac{2\pi d}{\lambda} \sin \theta, \quad \text{for } m = 0, 1, 2, \dots$$

or $d \sin \theta = m\lambda, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{maxima}), \quad (35-25)$

which is exactly Eq. 35-14, the expression that we derived earlier for the locations of the maxima.

The minima in the fringe pattern occur when

$$\frac{1}{2}\phi = (m + \frac{1}{2})\pi, \quad \text{for } m = 0, 1, 2, \dots \quad (35-26)$$

If we combine this relation with Eq. 35-23, we are led at once to

$$d \sin \theta = (m + \frac{1}{2})\lambda, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{minima}), \quad (35-27)$$

which is just Eq. 35-16, the expression we derived earlier for the locations of the fringe minima.

Figure 35-12, which is a plot of Eq. 35-22, shows the intensity of double-slit interference patterns as a function of the phase difference ϕ between the waves at the screen. The horizontal solid line is I_0 , the (uniform) intensity on the screen when one of the slits is covered up. Note in Eq. 35-22 and the graph that the intensity I varies from zero at the fringe minima to $4I_0$ at the fringe maxima.

If the waves from the two sources (slits) were *incoherent*, so that no enduring phase relation existed between them, there would be no fringe pattern and the intensity would have the uniform value $2I_0$ for all points on the screen; the horizontal dashed line in Fig. 35-12 shows this uniform value.

Interference cannot create or destroy energy but merely redistributes it over the screen. Thus, the *average* intensity on the screen must be the same $2I_0$ regardless of whether the sources are coherent. This follows at once from Eq. 35-22; if we substitute $\frac{1}{2}$, the average value of the cosine-squared function, this equation reduces to $I_{\text{avg}} = 2I_0$.

Proof of Eqs. 35-22 and 35-23

We shall combine the electric field components E_1 and E_2 , given by Eqs. 35-20 and 35-21, respectively, by the method of phasors as is discussed in Section 16-11. In Fig. 35-13a, the waves with components E_1 and E_2 are represented by phasors of magnitude E_0 that rotate around the origin at angular speed ω . The values of E_1 and E_2 at any time are the projections of the corresponding phasors on the vertical axis. Figure 35-13a shows the phasors and their projections at an arbitrary time t . Consistent with Eqs. 35-20 and 35-21, the phasor for E_1 has a rotation angle ωt and the phasor for E_2 has a rotation angle $\omega t + \phi$ (it is phase-shifted ahead of E_1). As each phasor rotates, its projection on the vertical axis varies with time in the same way that the sinusoidal functions of Eqs. 35-20 and 35-21 vary with time.

To combine the field components E_1 and E_2 at any point P in Fig. 35-10, we add their phasors vectorially, as shown in Fig. 35-13b. The magnitude of the vector sum is the amplitude E of the resultant wave at point P , and that wave has a certain phase constant β . To find the amplitude E in Fig. 35-13b, we first note that the two angles marked β are equal because they are opposite equal-length sides of a triangle. From the theorem (for triangles) that an exterior angle (here ϕ , as shown in Fig. 35-13b) is equal to the sum of the two opposite interior angles (here that sum is $\beta + \beta$), we see that $\beta = \frac{1}{2}\phi$. Thus, we have

$$\begin{aligned} E &= 2(E_0 \cos \beta) \\ &= 2E_0 \cos \frac{1}{2}\phi. \end{aligned} \quad (35-28)$$

If we square each side of this relation, we obtain

$$E^2 = 4E_0^2 \cos^2 \frac{1}{2}\phi. \quad (35-29)$$

Now, from Eq. 33-24, we know that the intensity of an electromagnetic wave is proportional to the square of its amplitude. Therefore, the waves we are combining in Fig. 35-13b, whose amplitudes are E_0 , each has an intensity I_0 that is proportional to E_0^2 , and the resultant wave, with amplitude E , has an intensity I that is proportional to E^2 . Thus,

$$\frac{I}{I_0} = \frac{E^2}{E_0^2}.$$

Substituting Eq. 35-29 into this equation and rearranging then yield

$$I = 4I_0 \cos^2 \frac{1}{2}\phi,$$

which is Eq. 35-22, which we set out to prove.

We still must prove Eq. 35-23, which relates the phase difference ϕ between the waves arriving at any point P on the screen of Fig. 35-10 to the angle θ that serves as a locator of that point.

The phase difference ϕ in Eq. 35-21 is associated with the path length difference $S_1 b$ in Fig. 35-10b. If $S_1 b$ is $\frac{1}{2}\lambda$, then ϕ is π ; if $S_1 b$ is λ , then ϕ is 2π , and so on. This suggests

$$\left(\frac{\text{phase}}{\text{difference}} \right) = \frac{2\pi}{\lambda} \left(\frac{\text{path length}}{\text{difference}} \right). \quad (35-30)$$

The path length difference $S_1 b$ in Fig. 35-10b is $d \sin \theta$; so Eq. 35-30 for the phase difference between the two waves arriving at point P on the screen becomes

$$\phi = \frac{2\pi d}{\lambda} \sin \theta,$$

which is Eq. 35-23, the other equation that we set out to prove to relate ϕ to the angle θ that locates P .

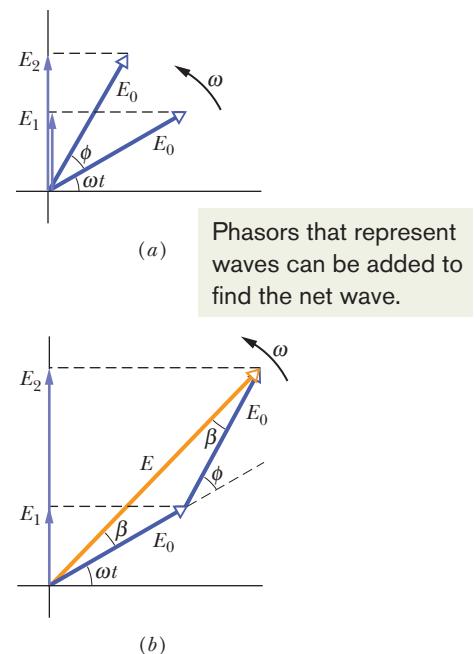


Fig. 35-13 (a) Phasors representing, at time t , the electric field components given by Eqs. 35-20 and 35-21. Both phasors have magnitude E_0 and rotate with angular speed ω . Their phase difference is ϕ . (b) Vector addition of the two phasors gives the phasor representing the resultant wave, with amplitude E and phase constant β .

Combining More Than Two Waves

In a more general case, we might want to find the resultant of more than two sinusoidally varying waves at a point. Whatever the number of waves is, our general procedure is this:

1. Construct a series of phasors representing the waves to be combined. Draw them end to end, maintaining the proper phase relations between adjacent phasors.
2. Construct the vector sum of this array. The length of this vector sum gives the amplitude of the resultant phasor. The angle between the vector sum and the first phasor is the phase of the resultant with respect to this first phasor. The projection of this vector-sum phasor on the vertical axis gives the time variation of the resultant wave.

CHECKPOINT 4

Each of four pairs of light waves arrives at a certain point on a screen. The waves have the same wavelength. At the arrival point, their amplitudes and phase differences are (a) $2E_0$, $6E_0$, and π rad; (b) $3E_0$, $5E_0$, and π rad; (c) $9E_0$, $7E_0$, and 3π rad; (d) $2E_0$, $2E_0$, and 0 rad. Rank the four pairs according to the intensity of the light at the arrival point, greatest first. (*Hint:* Draw phasors.)

Sample Problem

Combining three light waves by using phasors

Three light waves combine at a certain point where their electric field components are

$$\begin{aligned} E_1 &= E_0 \sin \omega t, \\ E_2 &= E_0 \sin(\omega t + 60^\circ), \\ E_3 &= E_0 \sin(\omega t - 30^\circ). \end{aligned}$$

Find their resultant component $E(t)$ at that point.

KEY IDEA

The resultant wave is

$$E(t) = E_1(t) + E_2(t) + E_3(t).$$

We can use the method of phasors to find this sum, and we are free to evaluate the phasors at any time t .

Calculations: To simplify the solution, we choose $t = 0$, for which the phasors representing the three waves are shown in Fig. 35-14. We can add these three phasors either directly on a vector-capable calculator or by components. For the component approach, we first write the sum of their horizontal components as

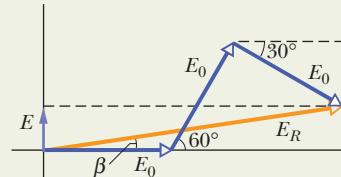
$$\sum E_h = E_0 \cos 0 + E_0 \cos 60^\circ + E_0 \cos(-30^\circ) = 2.37E_0.$$

The sum of their vertical components, which is the value of E at $t = 0$, is

$$\sum E_v = E_0 \sin 0 + E_0 \sin 60^\circ + E_0 \sin(-30^\circ) = 0.366E_0.$$

Fig. 35-14 Three phasors, representing waves with equal amplitudes E_0 and with phase constants 0° , 60° , and -30° , shown at time $t = 0$. The phasors combine to give a resultant phasor with magnitude E_R , at angle β .

Phasors that represent waves can be added to find the net wave.



The resultant wave $E(t)$ thus has an amplitude E_R of

$$E_R = \sqrt{(2.37E_0)^2 + (0.366E_0)^2} = 2.4E_0,$$

and a phase angle β relative to the phasor representing E_1 of

$$\beta = \tan^{-1}\left(\frac{0.366E_0}{2.37E_0}\right) = 8.8^\circ.$$

We can now write, for the resultant wave $E(t)$,

$$\begin{aligned} E &= E_R \sin(\omega t + \beta) \\ &= 2.4E_0 \sin(\omega t + 8.8^\circ). \end{aligned} \quad (\text{Answer})$$

Be careful to interpret the angle β correctly in Fig. 35-14: It is the constant angle between E_R and the phasor representing E_1 as the four phasors rotate as a single unit around the origin. The angle between E_R and the horizontal axis in Fig. 35-14 does not remain equal to β .



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35-7 Interference from Thin Films

The colors we see when sunlight illuminates a soap bubble or an oil slick are caused by the interference of light waves reflected from the front and back surfaces of a thin transparent film. The thickness of the soap or oil film is typically of the order of magnitude of the wavelength of the (visible) light involved. (Greater thicknesses spoil the coherence of the light needed to produce the colors due to interference.)

Figure 35-15 shows a thin transparent film of uniform thickness L and index of refraction n_2 , illuminated by bright light of wavelength λ from a distant point source. For now, we assume that air lies on both sides of the film and thus that $n_1 = n_3$ in Fig. 35-15. For simplicity, we also assume that the light rays are almost perpendicular to the film ($\theta \approx 0$). We are interested in whether the film is bright or dark to an observer viewing it almost perpendicularly. (Since the film is brightly illuminated, how could it possibly be dark? You will see.)

The incident light, represented by ray i , intercepts the front (left) surface of the film at point a and undergoes both reflection and refraction there. The reflected ray r_1 is intercepted by the observer's eye. The refracted light crosses the film to point b on the back surface, where it undergoes both reflection and refraction. The light reflected at b crosses back through the film to point c , where it undergoes both reflection and refraction. The light refracted at c , represented by ray r_2 , is intercepted by the observer's eye.

If the light waves of rays r_1 and r_2 are exactly in phase at the eye, they produce an interference maximum and region ac on the film is bright to the observer. If they are exactly out of phase, they produce an interference minimum and region ac is dark to the observer, *even though it is illuminated*. If there is some intermediate phase difference, there are intermediate interference and brightness.

Thus, the key to what the observer sees is the phase difference between the waves of rays r_1 and r_2 . Both rays are derived from the same ray i , but the path involved in producing r_2 involves light traveling twice across the film (a to b , and then b to c), whereas the path involved in producing r_1 involves no travel through the film. Because θ is about zero, we approximate the path length difference between the waves of r_1 and r_2 as $2L$. However, to find the phase difference between the waves, we cannot just find the number of wavelengths λ that is equivalent to a path length difference of $2L$. This simple approach is impossible

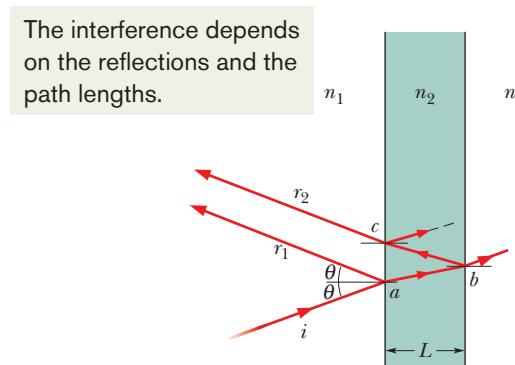


Fig. 35-15 Light waves, represented with ray i , are incident on a thin film of thickness L and index of refraction n_2 . Rays r_1 and r_2 represent light waves that have been reflected by the front and back surfaces of the film, respectively. (All three rays are actually nearly perpendicular to the film.) The interference of the waves of r_1 and r_2 with each other depends on their phase difference. The index of refraction n_1 of the medium at the left can differ from the index of refraction n_3 of the medium at the right, but for now we assume that both media are air, with $n_1 = n_3 = 1.0$, which is less than n_2 .

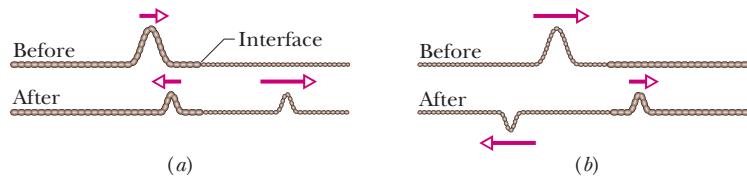


Fig. 35-16 Phase changes when a pulse is reflected at the interface between two stretched strings of different linear densities. The wave speed is greater in the lighter string. (a) The incident pulse is in the denser string. (b) The incident pulse is in the lighter string. Only here is there a phase change, and only in the reflected wave.

for two reasons: (1) the path length difference occurs in a medium other than air, and (2) reflections are involved, which can change the phase.



The phase difference between two waves can change if one or both are reflected.

Before we continue our discussion of interference from thin films, we must discuss changes in phase that are caused by reflections.

Reflection Phase Shifts

Refraction at an interface never causes a phase change—but reflection can, depending on the indexes of refraction on the two sides of the interface. Figure 35-16 shows what happens when reflection causes a phase change, using as an example pulses on a denser string (along which pulse travel is relatively slow) and a lighter string (along which pulse travel is relatively fast).

When a pulse traveling relatively slowly along the denser string in Fig. 35-16a reaches the interface with the lighter string, the pulse is partially transmitted and partially reflected, with no change in orientation. For light, this situation corresponds to the incident wave traveling in the medium of greater index of refraction n (recall that greater n means slower speed). In that case, the wave that is reflected at the interface does not undergo a change in phase; that is, its *reflection phase shift* is zero.

When a pulse traveling more quickly along the lighter string in Fig. 35-16b reaches the interface with the denser string, the pulse is again partially transmitted and partially reflected. The transmitted pulse again has the same orientation as the incident pulse, but now the reflected pulse is inverted. For a sinusoidal wave, such an inversion involves a phase change of π rad, or half a wavelength. For light, this situation corresponds to the incident wave traveling in the medium of lesser index of refraction (with greater speed). In that case, the wave that is reflected at the interface undergoes a phase shift of π rad, or half a wavelength.

We can summarize these results for light in terms of the index of refraction of the medium off which (or from which) the light reflects:



	Reflection	Reflection phase shift
Off lower index		0
Off higher index		0.5 wavelength

This might be remembered as “higher means half.”

Equations for Thin-Film Interference

In this chapter we have now seen three ways in which the phase difference between two waves can change:

35-7 INTERFERENCE FROM THIN FILMS

975

1. by reflection
2. by the waves traveling along paths of different lengths
3. by the waves traveling through media of different indexes of refraction

When light reflects from a thin film, producing the waves of rays r_1 and r_2 shown in Fig. 35-15, all three ways are involved. Let us consider them one by one.

We first reexamine the two reflections in Fig. 35-15. At point a on the front interface, the incident wave (in air) reflects from the medium having the higher of the two indexes of refraction; so the wave of reflected ray r_1 has its phase shifted by 0.5 wavelength. At point b on the back interface, the incident wave reflects from the medium (air) having the lower of the two indexes of refraction; so the wave reflected there is not shifted in phase by the reflection, and thus neither is the portion of it that exits the film as ray r_2 . We can organize this information with the first line in Table 35-1, which refers to the simplified drawing in Fig. 35-17 for a thin film in air. So far, as a result of the reflection phase shifts, the waves of r_1 and r_2 have a phase difference of 0.5 wavelength and thus are exactly out of phase.

Now we must consider the path length difference $2L$ that occurs because the wave of ray r_2 crosses the film twice. (This difference $2L$ is shown on the second line in Table 35-1.) If the waves of r_1 and r_2 are to be exactly in phase so that they produce fully constructive interference, the path length $2L$ must cause an additional phase difference of 0.5, 1.5, 2.5, . . . wavelengths. Only then will the net phase difference be an integer number of wavelengths. Thus, for a bright film, we must have

$$2L = \frac{\text{odd number}}{2} \times \text{wavelength} \quad (\text{in-phase waves}). \quad (35-31)$$

The wavelength we need here is the wavelength λ_{n2} of the light in the medium containing path length $2L$ —that is, in the medium with index of refraction n_2 . Thus, we can rewrite Eq. 35-31 as

$$2L = \frac{\text{odd number}}{2} \times \lambda_{n2} \quad (\text{in-phase waves}). \quad (35-32)$$

If, instead, the waves are to be exactly out of phase so that there is fully destructive interference, the path length $2L$ must cause either no additional phase difference or a phase difference of 1, 2, 3, . . . wavelengths. Only then will the net phase difference be an odd number of half-wavelengths. For a dark film, we must have

$$2L = \text{integer} \times \text{wavelength} \quad (\text{out-of-phase waves}). \quad (35-33)$$

where, again, the wavelength is the wavelength λ_{n2} in the medium containing $2L$. Thus, this time we have

$$2L = \text{integer} \times \lambda_{n2} \quad (\text{out-of-phase waves}). \quad (35-34)$$

Now we can use Eq. 35-8 ($\lambda_n = \lambda/n$) to write the wavelength of the wave of ray r_2 inside the film as

$$\lambda_{n2} = \frac{\lambda}{n_2}, \quad (35-35)$$

where λ is the wavelength of the incident light in vacuum (and approximately also in air). Substituting Eq. 35-35 into Eq. 35-32 and replacing “odd number/2” with $(m + \frac{1}{2})$ give us

$$2L = (m + \frac{1}{2}) \frac{\lambda}{n_2}, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{maxima—bright film in air}). \quad (35-36)$$

Table 35-1

An Organizing Table for Thin-Film Interference in Air (Fig. 35-17)^a

	r_1	r_2
Reflection phase shifts	0.5 wavelength	0
Path length difference		$2L$
Index in which path length difference occurs		n_2
In phase ^a : $2L = \frac{\text{odd number}}{2} \times \frac{\lambda}{n_2}$		
Out of phase ^a : $2L = \text{integer} \times \frac{\lambda}{n_2}$		

^aValid for $n_2 > n_1$ and $n_2 > n_3$.

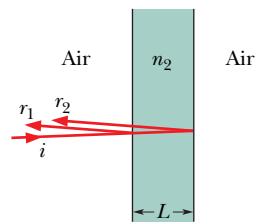


Fig. 35-17 Reflections from a thin film in air.



Fig. 35-18 The reflection of light from a soapy water film spanning a vertical loop. The top portion is so thin that the light reflected there undergoes destructive interference, making that portion dark. Colored interference fringes, or bands, decorate the rest of the film but are marred by circulation of liquid within the film as the liquid is gradually pulled downward by gravitation. (Richard Megna/Fundamental Photographs)

Similarly, with m replacing “integer,” Eq. 35-34 yields

$$2L = m \frac{\lambda}{n_2}, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{minima—dark film in air}). \quad (35-37)$$

For a given film thickness L , Eqs. 35-36 and 35-37 tell us the wavelengths of light for which the film appears bright and dark, respectively, one wavelength for each value of m . Intermediate wavelengths give intermediate brightnesses. For a given wavelength λ , Eqs. 35-36 and 35-37 tell us the thicknesses of the films that appear bright and dark in that light, respectively, one thickness for each value of m . Intermediate thicknesses give intermediate brightnesses.

Film Thickness Much Less Than λ

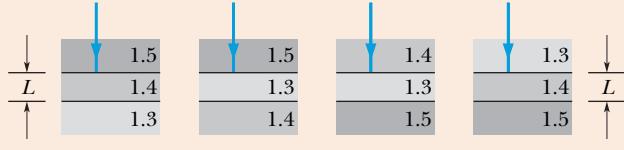
A special situation arises when a film is so thin that L is much less than λ , say, $L < 0.1\lambda$. Then the path length difference $2L$ can be neglected, and the phase difference between r_1 and r_2 is due *only* to reflection phase shifts. If the film of Fig. 35-17, where the reflections cause a phase difference of 0.5 wavelength, has thickness $L < 0.1\lambda$, then r_1 and r_2 are exactly out of phase, and thus the film is dark, regardless of the wavelength and intensity of the light. This special situation corresponds to $m = 0$ in Eq. 35-37. We shall count *any* thickness $L < 0.1\lambda$ as being the least thickness specified by Eq. 35-37 to make the film of Fig. 35-17 dark. (Every such thickness will correspond to $m = 0$.) The next greater thickness that will make the film dark is that corresponding to $m = 1$.

Figure 35-18 shows a vertical soap film whose thickness increases from top to bottom because gravitation has caused the film to slump. Bright white light illuminates the film. However, the top portion is so thin that it is dark. In the (somewhat thicker) middle we see fringes, or bands, whose color depends primarily on the wavelength at which reflected light undergoes fully constructive interference for a particular thickness. Toward the (thickest) bottom the fringes become progressively narrower and the colors begin to overlap and fade.



CHECKPOINT 5

The figure shows four situations in which light reflects perpendicularly from a thin film of thickness L , with indexes of refraction as given. (a) For which situations does reflection at the film interfaces cause a zero phase difference for the two reflected rays? (b) For which situations will the film be dark if the path length difference $2L$ causes a phase difference of 0.5 wavelength?



Sample Problem

Thin-film interference of a water film in air

White light, with a uniform intensity across the visible wavelength range of 400 to 690 nm, is perpendicularly incident on a water film, of index of refraction $n_2 = 1.33$ and thickness $L = 320$ nm, that is suspended in air. At what wavelength λ is the light reflected by the film brightest to an observer?

KEY IDEA

The reflected light from the film is brightest at the wavelengths λ for which the reflected rays are in phase with one another. The equation relating these wavelengths λ to the given film thickness L and film index of refraction n_2 is either

Eq. 35-36 or Eq. 35-37, depending on the reflection phase shifts for this particular film.

Calculations: To determine which equation is needed, we should fill out an organizing table like Table 35-1. However, because there is air on both sides of the water film, the situation here is exactly like that in Fig. 35-17, and thus the table would be exactly like Table 35-1. Then from Table 35-1, we see that the reflected rays are in phase (and thus the film is brightest) when

$$2L = \frac{\text{odd number}}{2} \times \frac{\lambda}{n_2},$$

which leads to Eq. 35-36:

$$2L = (m + \frac{1}{2}) \frac{\lambda}{n_2}.$$

Solving for λ and substituting for L and n_2 , we find

$$\lambda = \frac{2n_2 L}{m + \frac{1}{2}} = \frac{(2)(1.33)(320 \text{ nm})}{m + \frac{1}{2}} = \frac{851 \text{ nm}}{m + \frac{1}{2}}.$$

For $m = 0$, this gives us $\lambda = 1700 \text{ nm}$, which is in the infrared region. For $m = 1$, we find $\lambda = 567 \text{ nm}$, which is yellow-green light, near the middle of the visible spectrum. For $m = 2$, $\lambda = 340 \text{ nm}$, which is in the ultraviolet region. Thus, the wavelength at which the light seen by the observer is brightest is

$$\lambda = 567 \text{ nm.} \quad (\text{Answer})$$

Sample Problem

Thin-film interference of a coating on a glass lens

In Fig. 35-19, a glass lens is coated on one side with a thin film of magnesium fluoride (MgF_2) to reduce reflection from the lens surface. The index of refraction of MgF_2 is 1.38; that of the glass is 1.50. What is the least coating thickness that eliminates (via interference) the reflections at the middle of the visible spectrum ($\lambda = 550 \text{ nm}$)? Assume that the light is approximately perpendicular to the lens surface.

KEY IDEA

Reflection is eliminated if the film thickness L is such that light waves reflected from the two film interfaces are exactly out of phase. The equation relating L to the given wavelength λ and the index of refraction n_2 of the thin film is either Eq. 35-36 or Eq. 35-37, depending on the reflection phase shifts at the interfaces.

Calculations: To determine which equation is needed, we fill out an organizing table like Table 35-1. At the first interface, the incident light is in air, which has a lesser index of refraction than the MgF_2 (the thin film). Thus, we fill in 0.5 wavelength under r_1 in our organizing table (meaning that the waves of ray r_1 are shifted by 0.5λ at the first interface). At the second interface, the incident light is in the MgF_2 , which has a lesser index of refraction than the glass on the other side of the interface. Thus, we fill in 0.5 wavelength under r_2 in our table.

Because both reflections cause the same phase shift, they tend to put the waves of r_1 and r_2 in phase. Since we want those waves to be *out of phase*, their path length difference $2L$ must be an odd number of half-wavelengths:

$$2L = \frac{\text{odd number}}{2} \times \frac{\lambda}{n_2}.$$

This leads to Eq. 35-36. Solving that equation for L then gives us the film thicknesses that will eliminate reflection from the lens and coating:

$$L = (m + \frac{1}{2}) \frac{\lambda}{2n_2}, \quad \text{for } m = 0, 1, 2, \dots \quad (35-38)$$

We want the least thickness for the coating—that is, the smallest value of L . Thus, we choose $m = 0$, the smallest possible value of m . Substituting it and the given data in Eq. 35-38, we obtain

$$L = \frac{\lambda}{4n_2} = \frac{550 \text{ nm}}{(4)(1.38)} = 99.6 \text{ nm.} \quad (\text{Answer})$$

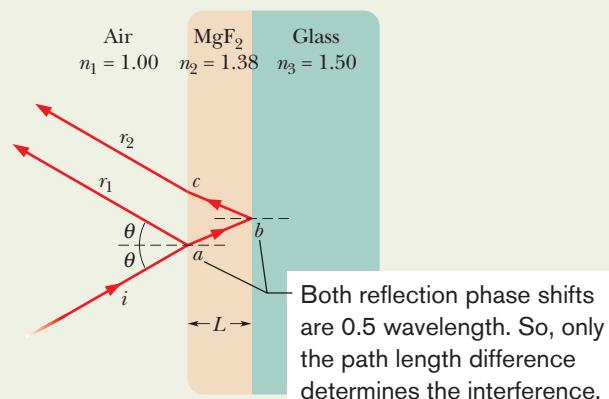


Fig. 35-19 Unwanted reflections from glass can be suppressed (at a chosen wavelength) by coating the glass with a thin transparent film of magnesium fluoride of the properly chosen thickness.

Sample Problem

Thin-film interference of a transparent wedge

Figure 35-20a shows a transparent plastic block with a thin wedge of air at the right. (The wedge thickness is exaggerated in the figure.) A broad beam of red light, with wavelength $\lambda = 632.8 \text{ nm}$, is directed downward through the top of the block (at an incidence angle of 0°). Some of the light that passes into the plastic is reflected back up from the top and bottom surfaces of the wedge, which acts as a thin film (of air) with a thickness that varies uniformly and gradually from L_L at the left-hand end to L_R at the right-hand end. (The plastic layers above and below the wedge of air are too thick to act as thin films.) An observer looking down on the block sees an interference pattern consisting of six dark fringes and five bright red fringes along the wedge. What is the change in thickness $\Delta L (= L_R - L_L)$ along the wedge?

KEY IDEAS

(1) The brightness at any point along the left-right length of the air wedge is due to the interference of the waves reflected at the top and bottom interfaces of the wedge. (2) The variation of brightness in the pattern of bright and dark fringes is due to the variation in the thickness of the wedge. In some regions, the thickness puts the reflected waves in phase and thus produces a bright reflection (a bright red fringe). In other regions, the thickness puts the reflected waves out of phase and thus produces no reflection (a dark fringe).

Organizing the reflections: Because the observer sees more dark fringes than bright fringes, we can assume that a dark fringe is produced at both the left and right ends of the wedge. Thus, the interference pattern is that shown in Fig. 35-20b.

We can represent the reflection of light at the top and bottom interfaces of the wedge, at any point along its length, with Fig. 35-20c, in which L is the wedge thickness at that point. Let us apply this figure to the left end of the wedge, where the reflections give a dark fringe.

We know that, for a dark fringe, the waves of rays r_1 and r_2 in Fig. 35-20e must be out of phase. We also know that the equation relating the film thickness L to the light's wavelength λ and the film's index of refraction n_2 is either Eq. 35-36 or Eq. 35-37, depending on the reflection phase shifts. To determine which equation gives a dark fringe at the left end of the wedge, we should fill out an organizing table like Table 35-1, as shown in Fig. 35-20e.

At the top interface of the wedge, the incident light is in the plastic, which has a greater n than the air beneath that interface. So, we fill in 0 under r_1 in our organizing table. At the bottom interface of the wedge, the incident light is in air, which has a lesser n than the plastic beneath that interface. So we fill in 0.5 wavelength under r_2 . So, the phase difference due to the reflection shifts is 0.5 wavelength. Thus the reflections alone tend to put the waves of r_1 and r_2 out of phase.

Reflections at left end (Fig. 35-20d): Because we see a dark fringe at the left end of the wedge, which the reflection phase shifts alone would produce, we don't want the path length difference to alter that condition. So, the path length difference $2L$ at the left end must be given by

$$2L = \text{integer} \times \frac{\lambda}{n_2},$$

which leads to Eq. 35-37:

$$2L = m \frac{\lambda}{n_2}, \quad \text{for } m = 0, 1, 2, \dots \quad (35-39)$$

Reflections at right end (Fig. 35-20f): Equation 35-39 holds not only for the left end of the wedge but also for any point along the wedge where a dark fringe is observed, including the right end, with a different integer value of m for each fringe. The least value of m is associated with the least thickness of the wedge where a dark fringe is observed. Progressively greater values of m are associated with progressively greater thicknesses of the wedge where a dark fringe is observed. Let m_L be the value at the left end. Then the value at the right end must be $m_L + 5$ because, from Fig. 35-20b, the right end is located at the fifth dark fringe from the left end.

Thickness difference: To find ΔL , we first solve Eq. 35-39 twice—once for the thickness L_L at the left end and once for the thickness L_R at the right end:

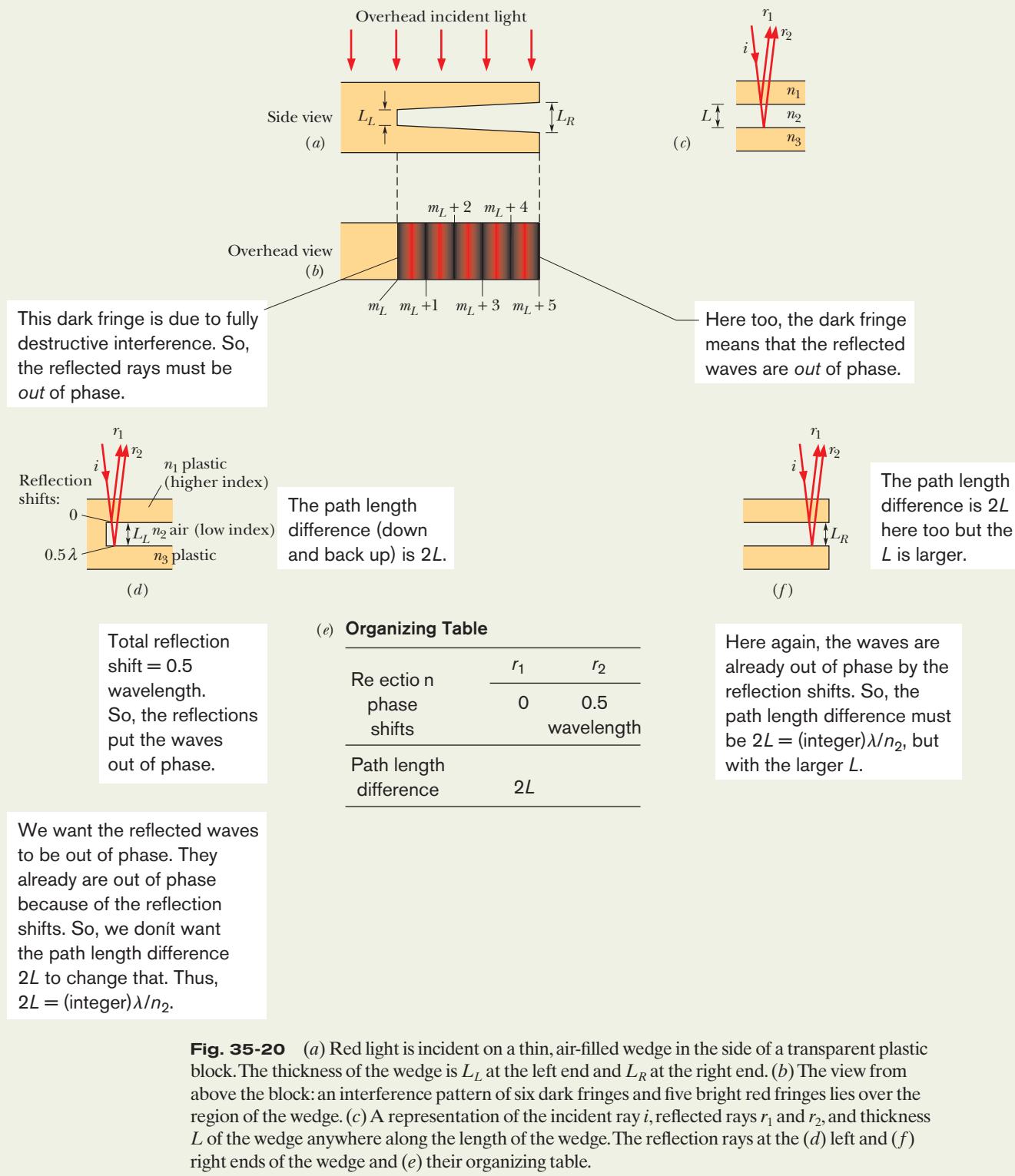
$$L_L = (m_L) \frac{\lambda}{2n_2}, \quad L_R = (m_L + 5) \frac{\lambda}{2n_2}. \quad (35-40)$$

We can now subtract L_L from L_R and substitute $n_2 = 1.00$ for the air within the wedge and $\lambda = 632.8 \times 10^{-9} \text{ m}$:

$$\begin{aligned} \Delta L = L_R - L_L &= \frac{(m_L + 5)\lambda}{2n_2} - \frac{m_L\lambda}{2n_2} = \frac{5}{2} \frac{\lambda}{n_2} \\ &= 1.58 \times 10^{-6} \text{ m}. \end{aligned} \quad (\text{Answer})$$



Additional examples, video, and practice available at WileyPLUS



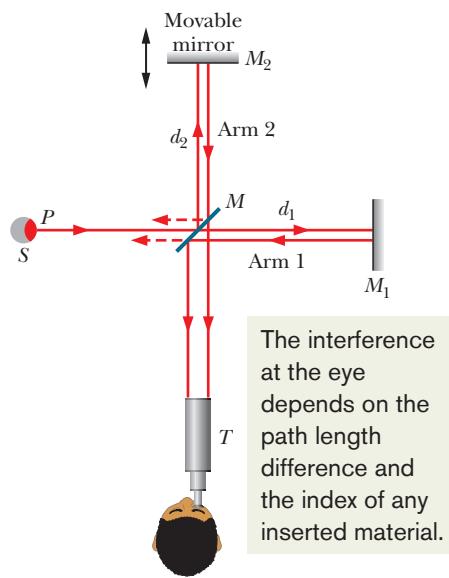


Fig. 35-21 Michelson's interferometer, showing the path of light originating at point P of an extended source S . Mirror M splits the light into two beams, which reflect from mirrors M_1 and M_2 back to M and then to telescope T . In the telescope an observer sees a pattern of interference fringes.

35-8 Michelson's Interferometer

An **interferometer** is a device that can be used to measure lengths or changes in length with great accuracy by means of interference fringes. We describe the form originally devised and built by A. A. Michelson in 1881.

Consider light that leaves point P on extended source S in Fig. 35-21 and encounters *beam splitter* M . A beam splitter is a mirror that transmits half the incident light and reflects the other half. In the figure we have assumed, for convenience, that this mirror possesses negligible thickness. At M the light thus divides into two waves. One proceeds by transmission toward mirror M_1 at the end of one arm of the instrument; the other proceeds by reflection toward mirror M_2 at the end of the other arm. The waves are entirely reflected at these mirrors and are sent back along their directions of incidence, each wave eventually entering telescope T . What the observer sees is a pattern of curved or approximately straight interference fringes; in the latter case the fringes resemble the stripes on a zebra.

The path length difference for the two waves when they recombine at the telescope is $2d_2 - 2d_1$, and anything that changes this path length difference will cause a change in the phase difference between these two waves at the eye. As an example, if mirror M_2 is moved by a distance $\frac{1}{2}\lambda$, the path length difference is changed by λ and the fringe pattern is shifted by one fringe (as if each dark stripe on a zebra had moved to where the adjacent dark stripe had been). Similarly, moving mirror M_2 by $\frac{1}{4}\lambda$ causes a shift by half a fringe (each dark zebra stripe shifts to where the adjacent white stripe had been).

A shift in the fringe pattern can also be caused by the insertion of a thin transparent material into the optical path of one of the mirrors—say, M_1 . If the material has thickness L and index of refraction n , then the number of wavelengths along the light's to-and-fro path through the material is, from Eq. 35-9,

$$N_m = \frac{2L}{\lambda_n} = \frac{2Ln}{\lambda}. \quad (35-41)$$

The number of wavelengths in the same thickness $2L$ of air before the insertion of the material is

$$N_a = \frac{2L}{\lambda}. \quad (35-42)$$

When the material is inserted, the light returned from mirror M_1 undergoes a phase change (in terms of wavelengths) of

$$N_m - N_a = \frac{2Ln}{\lambda} - \frac{2L}{\lambda} = \frac{2L}{\lambda}(n - 1). \quad (35-43)$$

For each phase change of one wavelength, the fringe pattern is shifted by one fringe. Thus, by counting the number of fringes through which the material causes the pattern to shift, and substituting that number for $N_m - N_a$ in Eq. 35-43, you can determine the thickness L of the material in terms of λ .

By such techniques the lengths of objects can be expressed in terms of the wavelengths of light. In Michelson's day, the standard of length—the meter—was the distance between two fine scratches on a certain metal bar preserved at Sèvres, near Paris. Michelson showed, using his interferometer, that the standard meter was equivalent to 1 553 163.5 wavelengths of a certain monochromatic red light emitted from a light source containing cadmium. For this careful measurement, Michelson received the 1907 Nobel Prize in physics. His work laid the foundation for the eventual abandonment (in 1961) of the meter bar as a standard of length and for the redefinition of the meter in terms of the wavelength of light. By 1983, even this wavelength standard was not precise enough to meet the growing requirements of science and technology, and it was replaced with a new standard based on a defined value for the speed of light.

REVIEW & SUMMARY

Huygens' Principle The three-dimensional transmission of waves, including light, may often be predicted by *Huygens' principle*, which states that all points on a wavefront serve as point sources of spherical secondary wavelets. After a time t , the new position of the wavefront will be that of a surface tangent to these secondary wavelets.

The law of refraction can be derived from Huygens' principle by assuming that the index of refraction of any medium is $n = c/v$, in which v is the speed of light in the medium and c is the speed of light in vacuum.

Wavelength and Index of Refraction The wavelength λ_n of light in a medium depends on the index of refraction n of the medium:

$$\lambda_n = \frac{\lambda}{n}, \quad (35-8)$$

in which λ is the wavelength in vacuum. Because of this dependency, the phase difference between two waves can change if they pass through different materials with different indexes of refraction.

Young's Experiment In *Young's interference experiment*, light passing through a single slit falls on two slits in a screen. The light leaving these slits flares out (by diffraction), and interference occurs in the region beyond the screen. A fringe pattern, due to the interference, forms on a viewing screen.

The light intensity at any point on the viewing screen depends in part on the difference in the path lengths from the slits to that point. If this difference is an integer number of wavelengths, the waves interfere constructively and an intensity maximum results. If it is an odd number of half-wavelengths, there is destructive interference and an intensity minimum occurs. The conditions for maximum and minimum intensity are

$$d \sin \theta = m\lambda, \quad \text{for } m = 0, 1, 2, \dots \\ (\text{maxima—bright fringes}), \quad (35-14)$$

$$d \sin \theta = (m + \frac{1}{2})\lambda, \quad \text{for } m = 0, 1, 2, \dots \\ (\text{minima—dark fringes}), \quad (35-16)$$

where θ is the angle the light path makes with a central axis and d is the slit separation.

Coherence If two light waves that meet at a point are to interfere perceptibly, the phase difference between them must remain constant with time; that is, the waves must be **coherent**. When two coherent waves meet, the resulting intensity may be found by using phasors.

Intensity in Two-Slit Interference In Young's interference experiment, two waves, each with intensity I_0 , yield a resultant wave of intensity I at the viewing screen, with

$$I = 4I_0 \cos^2 \frac{1}{2}\phi, \quad \text{where } \phi = \frac{2\pi d}{\lambda} \sin \theta. \quad (35-22, 35-23)$$

Equations 35-14 and 35-16, which identify the positions of the fringe maxima and minima, are contained within this relation.

Thin-Film Interference When light is incident on a thin transparent film, the light waves reflected from the front and back surfaces interfere. For near-normal incidence, the wavelength conditions for maximum and minimum intensity of the light reflected from a *film in air* are

$$2L = (m + \frac{1}{2}) \frac{\lambda}{n_2}, \quad \text{for } m = 0, 1, 2, \dots \\ (\text{maxima—bright film in air}), \quad (35-36)$$

$$2L = m \frac{\lambda}{n_2}, \quad \text{for } m = 0, 1, 2, \dots \\ (\text{minima—dark film in air}), \quad (35-37)$$

where n_2 is the index of refraction of the film, L is its thickness, and λ is the wavelength of the light in air.

If the light incident at an interface between media with different indexes of refraction is in the medium with the smaller index of refraction, the reflection causes a phase change of π rad, or half a wavelength, in the reflected wave. Otherwise, there is no phase change due to the reflection. Refraction causes no phase shift.

The Michelson Interferometer In *Michelson's interferometer* a light wave is split into two beams that, after traversing paths of different lengths, are recombined so they interfere and form a fringe pattern. Varying the path length of one of the beams allows distances to be accurately expressed in terms of wavelengths of light, by counting the number of fringes through which the pattern shifts because of the change.

QUESTIONS

1 Does the spacing between fringes in a two-slit interference pattern increase, decrease, or stay the same if (a) the slit separation is increased, (b) the color of the light is switched from red to blue, and (c) the whole apparatus is submerged in cooking sherry? (d) If the slits are illuminated with white light, then at any side maximum, does the blue component or the red component peak closer to the central maximum?

2 (a) If you move from one bright fringe in a two-slit interference pattern to the next one farther out, (b) does the path length difference ΔL increase or decrease and (c) by how much does it change, in wavelengths λ ?

3 Figure 35-22 shows two light rays that are initially exactly in phase and that reflect from several glass surfaces. Neglect the slight

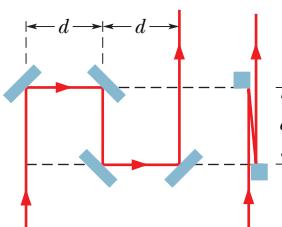


Fig. 35-22 Question 3.

982 CHAPTER 35 INTERFERENCE

slant in the path of the light in the second arrangement. (a) What is the path length difference of the rays? In wavelengths λ , (b) what should that path length difference equal if the rays are to be exactly out of phase when they emerge, and (c) what is the smallest value of d that will allow that final phase difference?

- 4 In Fig. 35-23, three pulses of light—*a*, *b*, and *c*—of the same wavelength are sent through layers of plastic having the given indexes of refraction and along the paths indicated. Rank the pulses according to their travel time through the plastic layers, greatest first.



Fig. 35-23 Question 4.

- 5 Is there an interference maximum, a minimum, an intermediate state closer to a maximum, or an intermediate state closer to a minimum at point *P* in Fig. 35-10 if the path length difference of the two rays is (a) 2.2λ , (b) 3.5λ , (c) 1.8λ , and (d) 1.0λ ? For each situation, give the value of m associated with the maximum or minimum involved.

- 6 Figure 35-24a gives intensity I versus position x on the viewing screen for the central portion of a two-slit interference pattern. The other parts of the figure give phasor diagrams for the electric field components of the waves arriving at the screen from the two slits (as in Fig. 35-13a). Which numbered points on the screen best correspond to which phasor diagram?

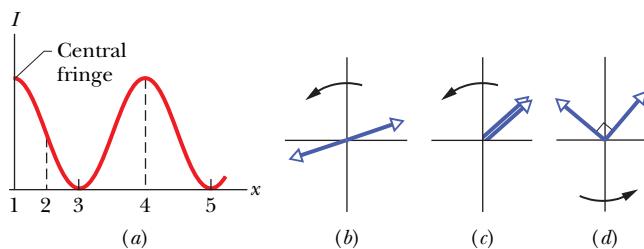


Fig. 35-24 Question 6.

- 7 Figure 35-25 shows two sources S_1 and S_2 that emit radio waves of wavelength λ in all directions. The sources are exactly in phase and are separated by a distance equal to 1.5λ . The vertical broken line is the perpendicular bisector of the distance between the sources. (a) If we start at the indicated start point and travel along path 1, does the interference produce a maximum all along the path, a minimum all along the path, or alternating maxima and minima? Repeat for (b) path 2 (along an axis through the sources) and (c) path 3 (along a perpendicular to that axis).

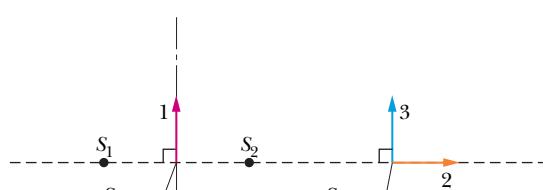


Fig. 35-25 Question 7.

- 8 Figure 35-26 shows two rays of light, of wavelength 600 nm, that reflect from glass surfaces separated by 150 nm. The rays are initially in phase. (a) What is the path length difference of the rays? (b) When they have cleared the reflection region, are the rays exactly in phase, exactly out of phase, or in some intermediate state?

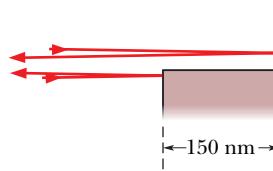


Fig. 35-26 Question 8.

- 9 Light travels along the length of a 1500-nm-long nanostructure. When a peak of the wave is at one end of the nanostructure, is there a peak or a valley at the other end if the wavelength is (a) 500 nm and (b) 1000 nm?

- 10 Figure 35-27a shows the cross section of a vertical thin film whose width increases downward because gravitation causes slumping. Figure 35-27b is a face-on view of the film, showing four bright (red) interference fringes that result when the film is illuminated with a perpendicular beam of red light. Points in the cross section corresponding to the bright fringes are labeled. In terms of the wavelength of the light inside the film, what is the difference in film thickness between (a) points *a* and *b* and (b) points *b* and *d*?

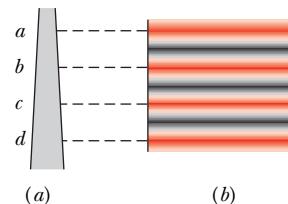


Fig. 35-27 Question 10.

- 11 Figure 35-28 shows four situations in which light reflects perpendicularly from a thin film of thickness L sandwiched between much thicker materials. The indexes of refraction are given. In which situations does Eq. 35-36 correspond to the reflections yielding maxima (that is, a bright film)?

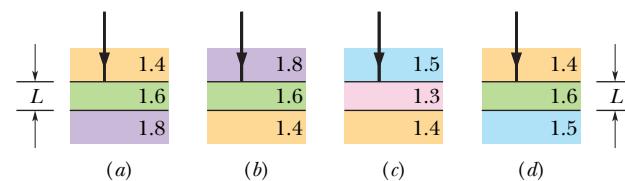


Fig. 35-28 Question 11.

- 12 Figure 35-29 shows the transmission of light through a thin film in air by a perpendicular beam (tilted in the figure for clarity). (a) Did ray r_3 undergo a phase shift due to reflection? (b) In wavelengths, what is the reflection phase shift for ray r_4 ? (c) If the film thickness is L , what is the path length difference between rays r_3 and r_4 ?

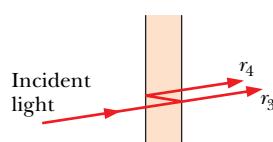


Fig. 35-29 Question 12.



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



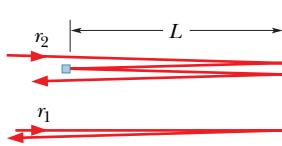
Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty

Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

PROBLEMS

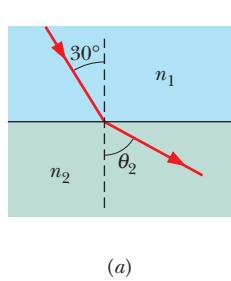
**Fig. 35-30**
Problems 1 and 2.**sec. 35-2 Light as a Wave**

- 1 In Fig. 35-30, a light wave along ray r_1 reflects once from a mirror and a light wave along ray r_2 reflects twice from that same mirror and once from a tiny mirror at distance L from the bigger mirror. (Neglect the slight tilt of the rays.) The waves have wavelength 620 nm and are initially in phase. (a) What is the smallest value of L that puts the final light waves exactly out of phase? (b) With the tiny mirror initially at that value of L , how far must it be moved away from the bigger mirror to again put the final waves out of phase?

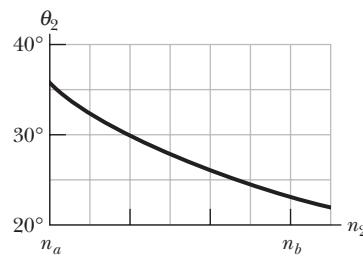
- 2 In Fig. 35-30, a light wave along ray r_1 reflects once from a mirror and a light wave along ray r_2 reflects twice from that same mirror and once from a tiny mirror at distance L from the bigger mirror. (Neglect the slight tilt of the rays.) The waves have wavelength λ and are initially exactly out of phase. What are the (a) smallest, (b) second smallest, and (c) third smallest values of L/λ that result in the final waves being exactly in phase?

- 3 **SSM** In Fig. 35-4, assume that two waves of light in air, of wavelength 400 nm, are initially in phase. One travels through a glass layer of index of refraction $n_1 = 1.60$ and thickness L . The other travels through an equally thick plastic layer of index of refraction $n_2 = 1.50$. (a) What is the smallest value L should have if the waves are to end up with a phase difference of 5.65 rad? (b) If the waves arrive at some common point with the same amplitude, is their interference fully constructive, fully destructive, intermediate but closer to fully constructive, or intermediate but closer to fully destructive?

- 4 In Fig. 35-31a, a beam of light in material 1 is incident on a boundary at an angle of 30° . The extent to which the light is bent due to refraction depends, in part, on the index of refraction n_2 of material 2. Figure 35-31b gives the angle of refraction θ_2 versus n_2 for a range of possible n_2 values, from $n_a = 1.30$ to $n_b = 1.90$. What is the speed of light in material 1?



(a)



(b)

Fig. 35-31 Problem 4.

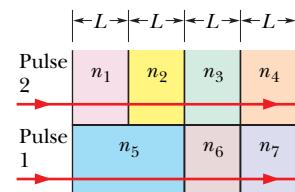
- 5 How much faster, in meters per second, does light travel in sapphire than in diamond? See Table 33-1.

- 6 The wavelength of yellow sodium light in air is 589 nm. (a) What is its frequency? (b) What is its wavelength in glass whose in-

dex of refraction is 1.52? (c) From the results of (a) and (b), find its speed in this glass.

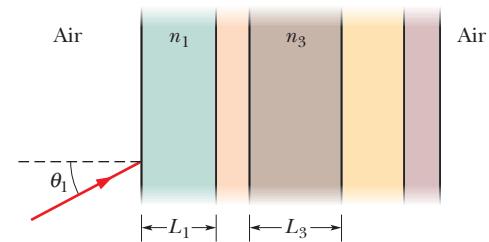
- 7 The speed of yellow light (from a sodium lamp) in a certain liquid is measured to be 1.92×10^8 m/s. What is the index of refraction of this liquid for the light?

- 8 In Fig. 35-32, two light pulses are sent through layers of plastic with thicknesses of either L or $2L$ as shown and indexes of refraction $n_1 = 1.55$, $n_2 = 1.70$, $n_3 = 1.60$, $n_4 = 1.45$, $n_5 = 1.59$, $n_6 = 1.65$, and $n_7 = 1.50$. (a) Which pulse travels through the plastic in less time? (b) What multiple of L/c gives the difference in the traversal times of the pulses?

**Fig. 35-32** Problem 8.

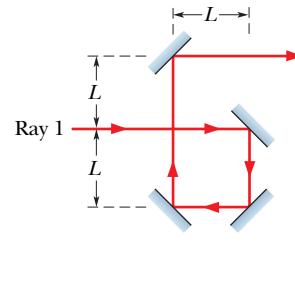
- 9 In Fig. 35-4, assume that the two light waves, of wavelength 620 nm in air, are initially out of phase by π rad. The indexes of refraction of the media are $n_1 = 1.45$ and $n_2 = 1.65$. What are the (a) smallest and (b) second smallest value of L that will put the waves exactly in phase once they pass through the two media?

- 10 In Fig. 35-33, a light ray is incident at angle $\theta_1 = 50^\circ$ on a series of five transparent layers with parallel boundaries. For layers 1 and 3, $L_1 = 20 \mu\text{m}$, $L_3 = 25 \mu\text{m}$, $n_1 = 1.6$, and $n_3 = 1.45$.

**Fig. 35-33** Problem 10.

- (a) At what angle does the light emerge back into air at the right? (b) How much time does the light take to travel through layer 3?

- 11 Suppose that the two waves in Fig. 35-4 have wavelength $\lambda = 500$ nm in air. What multiple of λ gives their phase difference when they emerge if (a) $n_1 = 1.50$, $n_2 = 1.60$, and $L = 8.50 \mu\text{m}$; (b) $n_1 = 1.62$, $n_2 = 1.72$, and $L = 8.50 \mu\text{m}$; and (c) $n_1 = 1.59$, $n_2 = 1.79$, and $L = 3.25 \mu\text{m}$? (d) Suppose that in each of these three situations the waves arrive at a common point (with the same amplitude) after emerging. Rank the situations according to the brightness the waves produce at the common point.

**Fig. 35-34** Problems 12 and 98.

- 12 In Fig. 35-34, two light rays go through different paths by reflecting from the various flat surfaces shown. The light waves have a wavelength of 420.0 nm and are initially in phase. What are the (a) smallest and (b) second smallest

value of distance L that will put the waves exactly out of phase as they emerge from the region?

- 13 GO ILW** Two waves of light in air, of wavelength $\lambda = 600.0 \text{ nm}$, are initially in phase. They then both travel through a layer of plastic as shown in Fig. 35-35, with $L_1 = 4.00 \mu\text{m}$, $L_2 = 3.50 \mu\text{m}$, $n_1 = 1.40$, and $n_2 = 1.60$. (a) What multiple of λ gives their phase difference after they both have emerged from the layers? (b) If the waves later arrive at some common point with the same amplitude, is their interference fully constructive, fully destructive, intermediate but closer to fully constructive, or intermediate but closer to fully destructive?

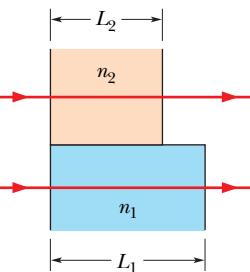


Fig. 35-35 Problem 13.

sec. 35-4 Young's Interference Experiment

- 14** In a double-slit arrangement the slits are separated by a distance equal to 100 times the wavelength of the light passing through the slits. (a) What is the angular separation in radians between the central maximum and an adjacent maximum? (b) What is the distance between these maxima on a screen 50.0 cm from the slits?

- 15 SSM** A double-slit arrangement produces interference fringes for sodium light ($\lambda = 589 \text{ nm}$) that have an angular separation of $3.50 \times 10^{-3} \text{ rad}$. For what wavelength would the angular separation be 10.0% greater?

- 16** A double-slit arrangement produces interference fringes for sodium light ($\lambda = 589 \text{ nm}$) that are 0.20° apart. What is the angular fringe separation if the entire arrangement is immersed in water ($n = 1.33$)?

- 17 GO SSM** In Fig. 35-36, two radio-frequency point sources S_1 and S_2 , separated by distance $d = 2.0 \text{ m}$, are radiating in phase with $\lambda = 0.50 \text{ m}$. A detector moves in a large circular path around the two sources in a plane containing them. How many maxima does it detect?

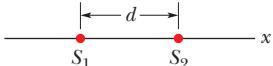


Fig. 35-36 Problems 17 and 22.

- 18** In the two-slit experiment of Fig. 35-10, let angle θ be 20.0° , the slit separation be $4.24 \mu\text{m}$, and the wavelength be $\lambda = 500 \text{ nm}$. (a) What multiple of λ gives the phase difference between the waves of rays r_1 and r_2 when they arrive at point P on the distant screen? (b) What is the phase difference in radians? (c) Determine where in the interference pattern point P lies by giving the maximum or minimum on which it lies, or the maximum and minimum between which it lies.

- 19 SSM ILW** Suppose that Young's experiment is performed with blue-green light of wavelength 500 nm. The slits are 1.20 mm apart, and the viewing screen is 5.40 m from the slits. How far apart are the bright fringes near the center of the interference pattern?

- 20** Monochromatic green light, of wavelength 550 nm, illuminates two parallel narrow slits $7.70 \mu\text{m}$ apart. Calculate the angular deviation (θ in Fig. 35-10) of the third-order ($m = 3$) bright fringe (a) in radians and (b) in degrees.

- 21** In a double-slit experiment, the distance between slits is 5.0 mm and the slits are 1.0 m from the screen. Two interference patterns can be seen on the screen: one due to light of wavelength 480 nm, and the other due to light of wavelength 600 nm. What is the

separation on the screen between the third-order ($m = 3$) bright fringes of the two interference patterns?

- 22** In Fig. 35-36, two isotropic point sources S_1 and S_2 emit identical light waves in phase at wavelength λ . The sources lie at separation d on an x axis, and a light detector is moved in a circle of large radius around the midpoint between them. It detects 30 points of zero intensity, including two on the x axis, one of them to the left of the sources and the other to the right of the sources. What is the value of d/λ ?

- 23** In Fig. 35-37, sources A and B emit long-range radio waves of wavelength 400 m, with the phase of the emission from A ahead of that from source B by 90° . The distance r_A from A to detector D is greater than the corresponding distance r_B by 100 m. What is the phase difference of the waves at D ?

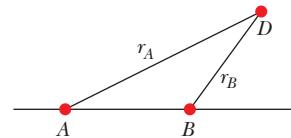


Fig. 35-37 Problem 23.

- 24** In Fig. 35-38, two isotropic point sources S_1 and S_2 emit light in phase at wavelength λ and at the same amplitude. The sources are separated by distance $2d = 6.00\lambda$. They lie on an axis that is parallel to an x axis, which runs along a viewing screen at distance $D = 20.0\lambda$. The origin lies on the perpendicular bisector between the sources. The figure shows two rays reaching point P on the screen, at position x_P . (a) At what value of x_P do the rays have the minimum possible phase difference? (b) What multiple of λ gives that minimum phase difference? (c) At what value of x_P do the rays have the maximum possible phase difference? What multiple of λ gives (d) that maximum phase difference and (e) the phase difference when $x_P = 6.00\lambda$? (f) When $x_P = 6.00\lambda$, is the resulting intensity at point P maximum, minimum, intermediate but closer to maximum, or intermediate but closer to minimum?

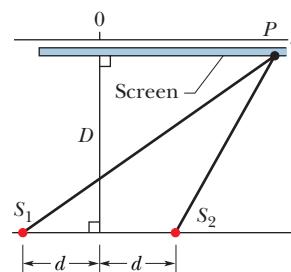


Fig. 35-38 Problem 24.

- 25 GO** In Fig. 35-39, two isotropic point sources of light (S_1 and S_2) are separated by distance $2.70 \mu\text{m}$ along a y axis and emit in phase at wavelength 900 nm and at the same amplitude. A light detector is located at point P at coordinate x_P on the x axis. What is the greatest value of x_P at which the detected light is minimum due to destructive interference?

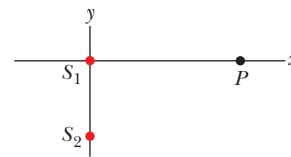


Fig. 35-39 Problems 25 and 28.

- 26** In a double-slit experiment, the fourth-order maximum for a wavelength of 450 nm occurs at an angle of $\theta = 90^\circ$. Thus, it is on the verge of being eliminated from the pattern because θ cannot exceed 90° in Eq. 35-14. (a) What range of wavelengths in the visible range (400 nm to 700 nm) are not present in the third-order maxima? To eliminate all of the visible light in the fourth-order maximum, (b) should the slit separation be increased or decreased and (c) what least change in separation is needed?

- 27** A thin flake of mica ($n = 1.58$) is used to cover one slit of a double-slit interference arrangement. The central point on the viewing screen is now occupied by what had been the seventh bright side fringe ($m = 7$). If $\lambda = 550 \text{ nm}$, what is the thickness of the mica?

••28 Figure 35-39 shows two isotropic point sources of light (S_1 and S_2) that emit in phase at wavelength 400 nm and at the same amplitude. A detection point P is shown on an x axis that extends through source S_1 . The phase difference ϕ between the light arriving at point P from the two sources is to be measured as P is moved along the x axis from $x = 0$ out to $x = +\infty$. The results out to $x_s = 10 \times 10^{-7}$ m are given in Fig. 35-40. On the way out to $+\infty$, what is the greatest value of x at which the light arriving at P from S_1 is exactly out of phase with the light arriving at P from S_2 ?

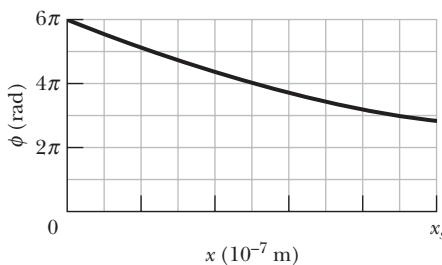


Fig. 35-40 Problem 28.

sec. 35-6 Intensity in Double-Slit Interference

•29 SSM Two waves of the same frequency have amplitudes 1.00 and 2.00. They interfere at a point where their phase difference is 60° . What is the resultant amplitude?

•30 Find the sum y of the following quantities:

$$y_1 = 10 \sin \omega t \quad \text{and} \quad y_2 = 8.0 \sin(\omega t + 30^\circ).$$

•31 ILW Add the quantities $y_1 = 10 \sin \omega t$, $y_2 = 15 \sin(\omega t + 30^\circ)$, and $y_3 = 5.0 \sin(\omega t - 45^\circ)$ using the phasor method.

•32 In the double-slit experiment of Fig. 35-10, the electric fields of the waves arriving at point P are given by

$$E_1 = (2.00 \mu\text{V/m}) \sin[(1.26 \times 10^{15})t]$$

$$E_2 = (2.00 \mu\text{V/m}) \sin[(1.26 \times 10^{15})t + 39.6 \text{ rad}],$$

where time t is in seconds. (a) What is the amplitude of the resultant electric field at point P ? (b) What is the ratio of the intensity I_P at point P to the intensity I_{cen} at the center of the interference pattern? (c) Describe where point P is in the interference pattern by giving the maximum or minimum on which it lies, or the maximum and minimum between which it lies. In a phasor diagram of the electric fields, (d) at what rate would the phasors rotate around the origin and (e) what is the angle between the phasors?

•33 GO Three electromagnetic waves travel through a certain point P along an x axis. They are polarized parallel to a y axis, with the following variations in their amplitudes. Find their resultant at P .

$$E_1 = (10.0 \mu\text{V/m}) \sin[(2.0 \times 10^{14} \text{ rad/s})t]$$

$$E_2 = (5.00 \mu\text{V/m}) \sin[(2.0 \times 10^{14} \text{ rad/s})t + 45.0^\circ]$$

$$E_3 = (5.00 \mu\text{V/m}) \sin[(2.0 \times 10^{14} \text{ rad/s})t - 45.0^\circ]$$

•34 In the double-slit experiment of Fig. 35-10, the viewing screen is at distance $D = 4.00 \text{ m}$, point P lies at distance $y = 20.5 \text{ cm}$ from the center of the pattern, the slit separation d is $4.50 \mu\text{m}$, and the wavelength λ is 580 nm. (a) Determine where point P is in the interference pattern by giving the maximum or minimum on which it lies, or the maximum and minimum between which it lies. (b) What is the ratio of the intensity I_P at point P to the intensity I_{cen} at the center of the pattern?

sec. 35-7 Interference from Thin Films

•35 SSM We wish to coat flat glass ($n = 1.50$) with a transparent material ($n = 1.25$) so that reflection of light at wavelength 600 nm is eliminated by interference. What minimum thickness can the coating have to do this?

•36 A 600-nm-thick soap film ($n = 1.40$) in air is illuminated with white light in a direction perpendicular to the film. For how many different wavelengths in the 300 to 700 nm range is there (a) fully constructive interference and (b) fully destructive interference in the reflected light?

•37 The rhinestones in costume jewelry are glass with index of refraction 1.50. To make them more reflective, they are often coated with a layer of silicon monoxide of index of refraction 2.00. What is the minimum coating thickness needed to ensure that light of wavelength 560 nm and of perpendicular incidence will be reflected from the two surfaces of the coating with fully constructive interference?

•38 White light is sent downward onto a horizontal thin film that is sandwiched between two materials. The indexes of refraction are 1.80 for the top material, 1.70 for the thin film, and 1.50 for the bottom material. The film thickness is $5.00 \times 10^{-7} \text{ m}$. Of the visible wavelengths (400 to 700 nm) that result in fully constructive interference at an observer above the film, which is the (a) longer and (b) shorter wavelength? The materials and film are then heated so that the film thickness increases. (c) Does the light resulting in fully constructive interference shift toward longer or shorter wavelengths?

•39 ILW Light of wavelength 624 nm is incident perpendicularly on a soap film ($n = 1.33$) suspended in air. What are the (a) least and (b) second least thicknesses of the film for which the reflections from the film undergo fully constructive interference?

•40 A thin film of acetone ($n = 1.25$) coats a thick glass plate ($n = 1.50$). White light is incident normal to the film. In the reflections, fully destructive interference occurs at 600 nm and fully constructive interference at 700 nm. Calculate the thickness of the acetone film.

•41 through 52 SSM 47, 51 GO

45, 49 *Reflection by thin layers*. In Fig. 35-41, light is incident perpendicularly on a thin layer of material 2 that lies between (thicker) materials 1 and 3. (The rays are tilted only for clarity.) The waves of rays r_1 and r_2 interfere, and here we consider the type of interference to be either maximum (max) or minimum (min). For this situation, each problem in Table 35-2 refers to the indexes of refraction n_1 , n_2 , and n_3 , the type of interference, the thin-layer thickness L in nanometers, and the wavelength λ in nanometers of the light as measured in air. Where λ is missing, give the wavelength that is in the visible range. Where L is missing, give the second least thickness or the third least thickness as indicated.

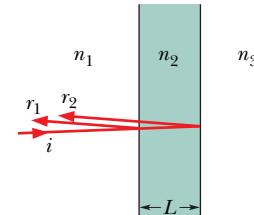


Fig. 35-41 Problems 41 through 52.

•53 The reflection of perpendicularly incident white light by a soap film in air has an interference maximum at 600 nm and a minimum at 450 nm, with no minimum in between. If $n = 1.33$ for the film, what is the film thickness, assumed uniform?

•54 A plane wave of monochromatic light is incident normally on a uniform thin film of oil that covers a glass plate. The wave-

length of the source can be varied continuously. Fully destructive interference of the reflected light is observed for wavelengths of 500 and 700 nm and for no wavelengths in between. If the index of refraction of the oil is 1.30 and that of the glass is 1.50, find the thickness of the oil film.

••55 SSM WWW A disabled tanker leaks kerosene ($n = 1.20$) into the Persian Gulf, creating a large slick on top of the water ($n = 1.30$). (a) If you are looking straight down from an airplane, while the Sun is overhead, at a region of the slick where its thickness is 460 nm, for which wavelength(s) of visible light is the reflection brightest because of constructive interference? (b) If you are scuba diving directly under this same region of the slick, for which wavelength(s) of visible light is the transmitted intensity strongest?

••56 A thin film, with a thickness of 272.7 nm and with air on both sides, is illuminated with a beam of white light. The beam is perpendicular to the film and consists of the full range of wavelengths for the visible spectrum. In the light reflected by the film, light with a wavelength of 600.0 nm undergoes fully constructive interference. At what wavelength does the reflected light undergo fully destructive interference? (Hint: You must make a reasonable assumption about the index of refraction.)

••57 through 68 SSM 59 GO 64, 65 *Transmission through thin layers.* In Fig. 35-42, light is incident perpendicularly on a thin layer of material 2 that lies between (thicker) materials 1 and 3. (The rays are tilted only for clarity.) Part of the light ends up in material 3 as ray r_3 (the light does not reflect inside material 2) and r_4 (the light reflects twice inside material 2). The waves of r_3 and r_4 interfere, and here we consider the type of interference to be either maximum (max) or minimum (min). For this situation, each problem in Table 35-3 refers to the indexes of refraction n_1 , n_2 , and n_3 , the type of interference, the thin-layer thickness L in nanometers, and the wavelength λ in nanometers of the light as measured in air. Where λ is missing, give the wavelength that is in the visible range. Where L is missing, give the second least thickness or the third least thickness as indicated.

••69 GO In Fig. 35-43, a broad beam of light of wavelength 630 nm is incident at 90° on a thin, wedge-shaped film with index of refraction 1.50. Transmission gives 10 bright and 9 dark fringes along the film's length. What is the left-to-right change in film thickness?

Table 35-2
Problems 41 through 52: Reflection by Thin Layers. See the setup for these problems.

	n_1	n_2	n_3	Type	L	λ
41	1.68	1.59	1.50	min	2nd	342
42	1.55	1.60	1.33	max	285	
43	1.60	1.40	1.80	min	200	
44	1.50	1.34	1.42	max	2nd	587
45	1.55	1.60	1.33	max	3rd	612
46	1.68	1.59	1.50	min	415	
47	1.50	1.34	1.42	min	380	
48	1.60	1.40	1.80	max	2nd	632
49	1.32	1.75	1.39	max	3rd	382
50	1.40	1.46	1.75	min	2nd	482
51	1.40	1.46	1.75	min	210	
52	1.32	1.75	1.39	max	325	

Table 35-3
Problems 57 through 68: Transmission Through Thin Layers.
See the setup for these problems.

	n_1	n_2	n_3	Type	L	λ
57	1.55	1.60	1.33	min	285	
58	1.32	1.75	1.39	min	3rd	382
59	1.68	1.59	1.50	max	415	
60	1.50	1.34	1.42	max	380	
61	1.32	1.75	1.39	min	325	
62	1.68	1.59	1.50	max	2nd	342
63	1.40	1.46	1.75	max	2nd	482
64	1.40	1.46	1.75	max	210	
65	1.60	1.40	1.80	min	2nd	632
66	1.60	1.40	1.80	max	200	
67	1.50	1.34	1.42	min	2nd	587
68	1.55	1.60	1.33	min	3rd	612

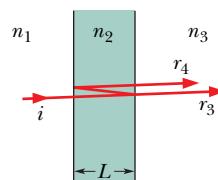


Fig. 35-42
Problems 57 through 68.

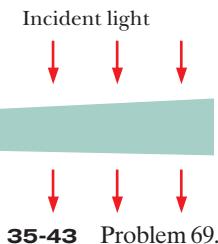


Fig. 35-43 Problem 69.

••70 In Fig. 35-44, a broad beam of light of wavelength 620 nm is sent directly downward through the top plate of a pair of glass plates touching at the left end. The air between the plates acts as a thin film, and an interference pattern can be seen from above the plates. Initially, a dark fringe lies at the left end, a bright fringe lies at the right end, and nine dark fringes lie between those two end fringes. The plates are then very gradually squeezed together at a constant rate to decrease the angle between them. As a result, the fringe at the right side changes between being bright to being dark every 15.0 s. (a) At what rate is the spacing between the plates at the right end being changed? (b) By how much has the spacing there changed when both left and right ends have a dark fringe and there are five dark fringes between them?

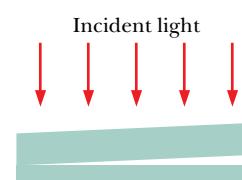


Fig. 35-44 Problems 70–74.

••71 In Fig. 35-44, two microscope slides touch at one end and are separated at the other end. When light of wavelength 500 nm

PROBLEMS

987

shines vertically down on the slides, an overhead observer sees an interference pattern on the slides with the dark fringes separated by 1.2 mm. What is the angle between the slides?

••72 In Fig. 35-44, a broad beam of monochromatic light is directed perpendicularly through two glass plates that are held together at one end to create a wedge of air between them. An observer intercepting light reflected from the wedge of air, which acts as a thin film, sees 4001 dark fringes along the length of the wedge. When the air between the plates is evacuated, only 4000 dark fringes are seen. Calculate to six significant figures the index of refraction of air from these data.

••73 SSM ILW In Fig. 35-44, a broad beam of light of wavelength 683 nm is sent directly downward through the top plate of a pair of glass plates. The plates are 120 mm long, touch at the left end, and are separated by 48.0 μm at the right end. The air between the plates acts as a thin film. How many bright fringes will be seen by an observer looking down through the top plate?

••74 GO Two rectangular glass plates ($n = 1.60$) are in contact along one edge and are separated along the opposite edge (Fig. 35-44). Light with a wavelength of 600 nm is incident perpendicularly onto the top plate. The air between the plates acts as a thin film. Nine dark fringes and eight bright fringes are observed from above the top plate. If the distance between the two plates along the separated edges is increased by 600 nm, how many dark fringes will there then be across the top plate?

••75 SSM ILW Figure 35-45a shows a lens with radius of curvature R lying on a flat glass plate and illuminated from above by light with wavelength λ . Figure 35-45b (a photograph taken from above the lens) shows that circular interference fringes (called Newton's rings) appear, associated with the variable thickness d of

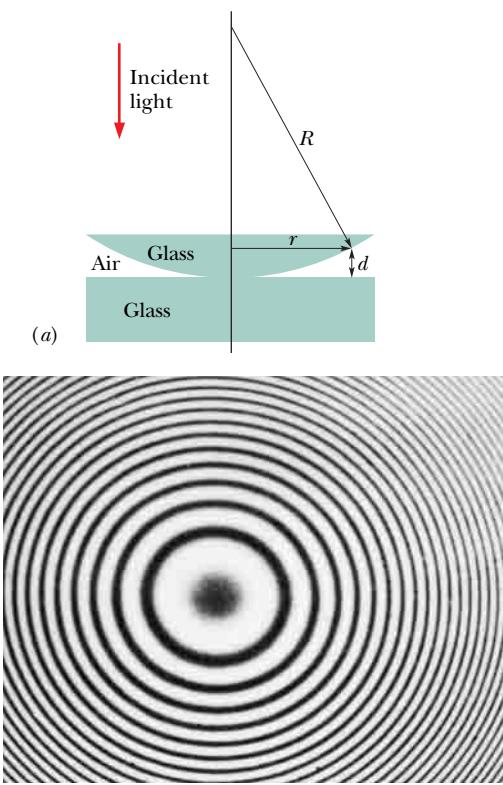


Fig. 35-45 Problems 75–77. (Courtesy Bausch & Lomb)

the air film between the lens and the plate. Find the radii r of the interference maxima assuming $r/R \ll 1$.

••76 In a Newton's rings experiment (see Problem 75), the radius of curvature R of the lens is 5.0 m and the lens diameter is 20 mm. (a) How many bright rings are produced? Assume that $\lambda = 589$ nm. (b) How many bright rings would be produced if the arrangement were immersed in water ($n = 1.33$)?

••77 A Newton's rings apparatus is to be used to determine the radius of curvature of a lens (see Fig. 35-45 and Problem 75). The radii of the n th and $(n + 20)$ th bright rings are measured and found to be 0.162 and 0.368 cm, respectively, in light of wavelength 546 nm. Calculate the radius of curvature of the lower surface of the lens.

••78 A thin film of liquid is held in a horizontal circular ring, with air on both sides of the film. A beam of light at wavelength 550 nm is directed perpendicularly onto the film, and the intensity I of its reflection is monitored. Figure 35-46 gives intensity I as a function of time t ; the horizontal scale is set by $t_s = 20.0$ s. The intensity changes because of evaporation from the two sides of the film. Assume that the film is flat and has parallel sides, a radius of 1.80 cm, and an index of refraction of 1.40. Also assume that the film's volume decreases at a constant rate. Find that rate.

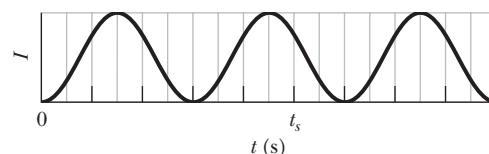


Fig. 35-46 Problem 78.

sec. 35-8 Michelson's Interferometer

••79 If mirror M_2 in a Michelson interferometer (Fig. 35-21) is moved through 0.233 mm, a shift of 792 bright fringes occurs. What is the wavelength of the light producing the fringe pattern?

••80 A thin film with index of refraction $n = 1.40$ is placed in one arm of a Michelson interferometer, perpendicular to the optical path. If this causes a shift of 7.0 bright fringes of the pattern produced by light of wavelength 589 nm, what is the film thickness?

••81 SSM WWW In Fig. 35-47, an airtight chamber of length $d = 5.0$ cm is placed in one of the arms of a Michelson interferometer. (The glass window on each end of the chamber has negligible thickness.) Light of wavelength $\lambda = 500$ nm is used. Evacuating the air from the chamber causes a shift of 60 bright fringes. From these data and to six significant figures, find the index of refraction of air at atmospheric pressure.

••82 The element sodium can emit light at two wavelengths, $\lambda_1 = 588.9950$ nm and $\lambda_2 = 589.5924$ nm. Light from sodium is being used in a Michelson interferometer (Fig. 35-21). Through what distance must mirror M_2 be moved if the shift in the fringe pattern for one wavelength is to be 1.00 fringe more than the shift in the fringe pattern for the other wavelength?

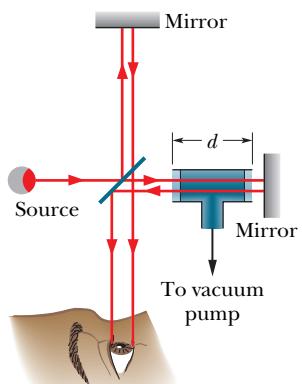


Fig. 35-47 Problem 81.

Additional Problems

- 83 GO** Two light rays, initially in phase and with a wavelength of 500 nm, go through different paths by reflecting from the various mirrors shown in Fig. 35-48. (Such a reflection does not itself produce a phase shift.) (a) What least value of distance d will put the rays exactly out of phase when they emerge from the region? (Ignore the slight tilt of the path for ray 2.) (b) Repeat the question assuming that the entire apparatus is immersed in a protein solution with an index of refraction of 1.38.

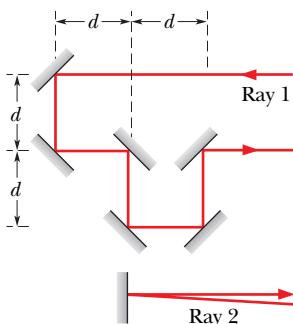


Fig. 35-48 Problem 83.

- 84** In Figure 35-49, two isotropic point sources S_1 and S_2 emit light in phase at wavelength λ and at the same amplitude. The sources are separated by distance $d = 6.00\lambda$ on an x axis. A viewing screen is at distance $D = 20.0\lambda$ from S_2 and parallel to the y axis. The figure shows two rays reaching point P on the screen, at height y_P . (a) At what value of y_P do the rays have the minimum possible phase difference? (b) What multiple of λ gives that minimum phase difference? (c) At what value of y_P do the rays have the maximum possible phase difference? What multiple of λ gives (d) that maximum phase difference and (e) the phase difference when $y_P = d$? (f) When $y_P = d$, is the resulting intensity at point P maximum, minimum, intermediate but closer to maximum, or intermediate but closer to minimum?

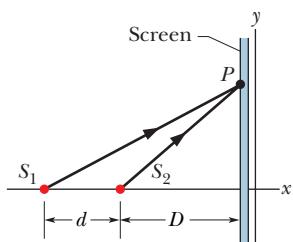


Fig. 35-49 Problem 84.

- 85 SSM** A double-slit arrangement produces bright interference fringes for sodium light (a distinct yellow light at a wavelength of $\lambda = 589$ nm). The fringes are angularly separated by 0.30° near the center of the pattern. What is the angular fringe separation if the entire arrangement is immersed in water, which has an index of refraction of 1.33?

- 86** In Fig. 35-50a, the waves along rays 1 and 2 are initially in phase, with the same wavelength λ in air. Ray 2 goes through a material with length L and index of refraction n . The rays are then reflected by mirrors to a common point P on a screen. Suppose that we can vary n from $n = 1.0$ to $n = 2.5$. Suppose also that, from $n = 1.0$ to $n = n_s = 1.5$, the intensity I of the light at point P varies with n as given in Fig. 35-50b. At what values of n greater than 1.4 is intensity I (a) maximum and (b) zero? (c) What multiple of λ gives the phase difference between the rays at point P when $n = 2.0$?

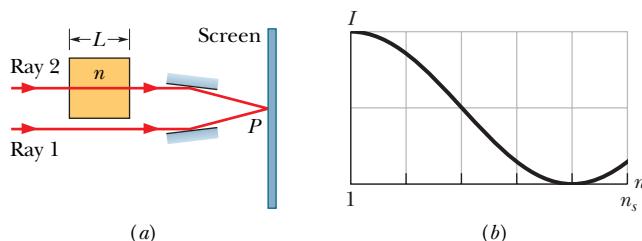


Fig. 35-50 Problems 86 and 87.

- 87 SSM** In Fig. 35-50a, the waves along rays 1 and 2 are initially in phase, with the same wavelength λ in air. Ray 2 goes through a material with length L and index of refraction n . The rays are then reflected by mirrors to a common point P on a screen. Suppose that we can vary L from 0 to 2400 nm. Suppose also that, from $L = 0$ to $L_s = 900$ nm, the intensity I of the light at point P varies with L as given in Fig. 35-51. At what values of L greater than L_s is intensity I (a) maximum and (b) zero? (c) What multiple of λ gives the phase difference between ray 1 and ray 2 at common point P when $L = 1200$ nm?

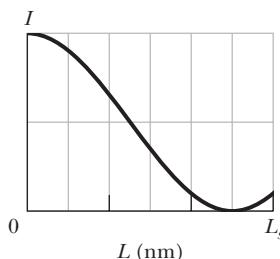


Fig. 35-51 Problem 87.

- 88** Light of wavelength 700.0 nm is sent along a route of length 2000 nm. The route is then filled with a medium having an index of refraction of 1.400. In degrees, by how much does the medium phase-shift the light? Give (a) the full shift and (b) the equivalent shift that has a value less than 360° .

- 89 SSM** In Fig. 35-52, a microwave transmitter at height a above the water level of a wide lake transmits microwaves of wavelength λ toward a receiver on the opposite shore, a distance x above the water level. The microwaves reflecting from the water interfere with the microwaves arriving directly from the transmitter. Assuming that the lake width D is much greater than a and x , and that $\lambda \geq a$, find an expression that gives the values of x for which the signal at the receiver is maximum. (Hint: Does the reflection cause a phase change?)

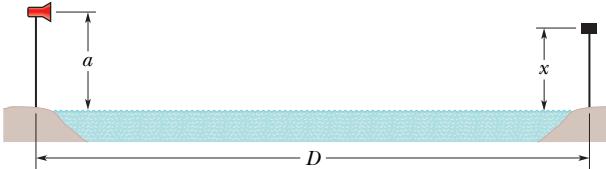


Fig. 35-52 Problem 89.

- 90** In Fig. 35-53, two isotropic point sources S_1 and S_2 emit light at wavelength $\lambda = 400$ nm. Source S_1 is located at $y = 640$ nm; source S_2 is located at $y = -640$ nm. At point P_1 (at $x = 720$ nm), the wave from S_2 arrives ahead of the wave from S_1 by a phase difference of 0.600π rad. (a) What multiple of λ gives the phase difference between the waves from the two sources as the waves arrive at point P_2 , which is located at $y = 720$ nm. (The figure is not drawn to scale.) (b) If the waves arrive at P_2 with equal amplitudes, is the interference there fully constructive, fully destructive, intermediate but closer to fully constructive, or intermediate but closer to fully destructive?

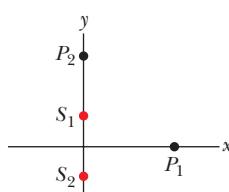


Fig. 35-53
Problem 90.

- 91** Ocean waves moving at a speed of 4.0 m/s are approaching a beach at angle $\theta_1 = 30^\circ$ to the normal, as shown from above in Fig. 35-54.

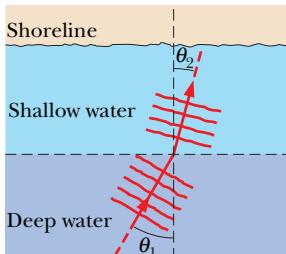


Fig. 35-54 Problem 91.

PROBLEMS

989

Suppose the water depth changes abruptly at a certain distance from the beach and the wave speed there drops to 3.0 m/s. (a) Close to the beach, what is the angle θ_2 between the direction of wave motion and the normal? (Assume the same law of refraction as for light.) (b) Explain why most waves come in normal to a shore even though at large distances they approach at a variety of angles.

92 Figure 35-55a shows two light rays that are initially in phase as they travel upward through a block of plastic, with wavelength 400 nm as measured in air. Light ray r_1 exits directly into air. However, before light ray r_2 exits into air, it travels through a liquid in a hollow cylinder within the plastic. Initially the height L_{liq} of the liquid is 40.0 μm , but then the liquid begins to evaporate. Let ϕ be the phase difference between rays r_1 and r_2 once they both exit into the air. Figure 35-55b shows ϕ versus the liquid's height L_{liq} until the liquid disappears, with ϕ given in terms of wavelength and the horizontal scale set by $L_s = 40.00 \mu\text{m}$. What are (a) the index of refraction of the plastic and (b) the index of refraction of the liquid?

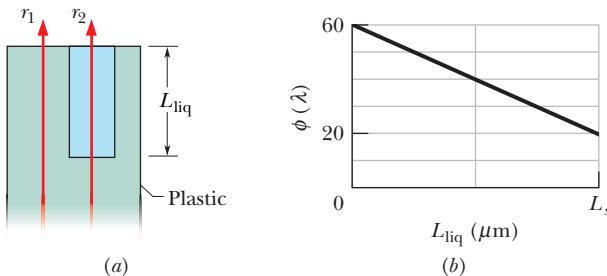


Fig. 35-55 Problem 92.

93 SSM If the distance between the first and tenth minima of a double-slit pattern is 18.0 mm and the slits are separated by 0.150 mm with the screen 50.0 cm from the slits, what is the wavelength of the light used?

94 Figure 35-56 shows an optical fiber in which a central plastic core of index of refraction $n_1 = 1.58$ is surrounded by a plastic sheath of index of refraction $n_2 = 1.53$. Light can travel along different paths within the central core, leading to different travel times through the fiber. This causes an initially short pulse of light to spread as it travels along the fiber, resulting in information loss. Consider light that travels directly along the central axis of the fiber and light that is repeatedly reflected at the critical angle along the core–sheath interface, reflecting from side to side as it travels down the central core. If the fiber length is 300 m, what is the difference in the travel times along these two routes?

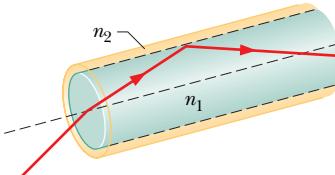


Fig. 35-56 Problem 94.

95 SSM Two parallel slits are illuminated with monochromatic light of wavelength 500 nm. An interference pattern is formed on a screen some distance from the slits, and the fourth dark band is located 1.68 cm from the central bright band on the screen. (a) What is the path length difference corresponding to the fourth dark band? (b) What is the distance on the screen between the central bright band and the first bright band on either side of the central band? (*Hint:* The angle to the fourth dark band and the angle to the first bright band are small enough that $\tan \theta \approx \sin \theta$.)

96 A camera lens with index of refraction greater than 1.30 is coated with a thin transparent film of index of refraction 1.25 to eliminate by interference the reflection of light at wavelength λ that is incident perpendicularly on the lens. What multiple of λ gives the minimum film thickness needed?

97 SSM Light of wavelength λ is used in a Michelson interferometer. Let x be the position of the movable mirror, with $x = 0$ when the arms have equal lengths $d_2 = d_1$. Write an expression for the intensity of the observed light as a function of x , letting I_m be the maximum intensity.

98 In two experiments, light is to be sent along the two paths shown in Fig. 35-34 by reflecting it from the various flat surfaces shown. In the first experiment, rays 1 and 2 are initially in phase and have a wavelength of 620.0 nm. In the second experiment, rays 1 and 2 are initially in phase and have a wavelength of 496.0 nm. What least value of distance L is required such that the 620.0 nm waves emerge from the region exactly in phase but the 496.0 nm waves emerge exactly out of phase?

99 Figure 35-57 shows the design of a Texas arcade game. Four laser pistols are pointed toward the center of an array of plastic layers where a clay armadillo is the target. The indexes of refraction of the layers are $n_1 = 1.55, n_2 = 1.70, n_3 = 1.45, n_4 = 1.60, n_5 = 1.45, n_6 = 1.61, n_7 = 1.59, n_8 = 1.70$, and $n_9 = 1.60$. The layer thicknesses are either 2.00 mm or 4.00 mm, as drawn. What is the travel time through the layers for the laser burst from (a) pistol 1, (b) pistol 2, (c) pistol 3, and (d) pistol 4? (e) If the pistols are fired simultaneously, which laser burst hits the target first?

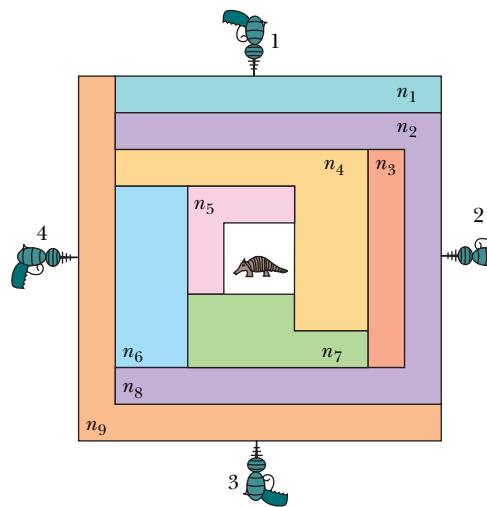


Fig. 35-57 Problem 99.

100 A thin film suspended in air is 0.410 μm thick and is illuminated with white light incident perpendicularly on its surface. The index of refraction of the film is 1.50. At what wavelength will visible light that is reflected from the two surfaces of the film undergo fully constructive interference?

101 Find the slit separation of a double-slit arrangement that will produce interference fringes 0.018 rad apart on a distant screen when the light has wavelength $\lambda = 589 \text{ nm}$.

102 In a phasor diagram for any point on the viewing screen for the two-slit experiment in Fig. 35-10, the resultant-wave phasor rotates 60.0° in $2.50 \times 10^{-16} \text{ s}$. What is the wavelength?

36

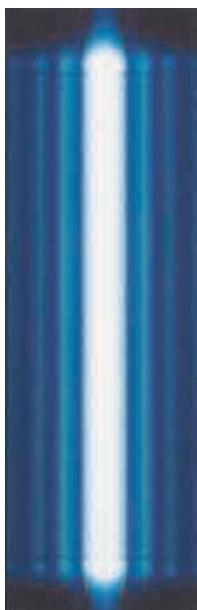


Fig. 36-1 This diffraction pattern appeared on a viewing screen when light that had passed through a narrow vertical slit reached the screen. Diffraction caused the light to flare out perpendicular to the long sides of the slit. That flaring produced an interference pattern consisting of a broad central maximum plus less intense and narrower secondary (or side) maxima, with minima between them. (Ken Kay/Fundamental Photographs)

DIFFRACTION

36-1

WHAT IS PHYSICS?

One focus of physics in the study of light is to understand and put to use the diffraction of light as it passes through a narrow slit or (as we shall discuss) past either a narrow obstacle or an edge. We touched on this phenomenon in Chapter 35 when we looked at how light flared—diffracted—through the slits in Young's experiment. Diffraction through a given slit is more complicated than simple flaring, however, because the light also interferes with itself and produces an interference pattern. It is because of such complications that light is rich with application opportunities. Even though the diffraction of light as it passes through a slit or past an obstacle seems awfully academic, countless engineers and scientists make their living using this physics, and the total worth of diffraction applications worldwide is probably incalculable.

Before we can discuss some of these applications, we first must discuss why diffraction is due to the wave nature of light.

36-2 Diffraction and the Wave Theory of Light

In Chapter 35 we defined diffraction rather loosely as the flaring of light as it emerges from a narrow slit. More than just flaring occurs, however, because the light produces an interference pattern called a **diffraction pattern**. For example, when monochromatic light from a distant source (or a laser) passes through a narrow slit and is then intercepted by a viewing screen, the light produces on the screen a diffraction pattern like that in Fig. 36-1. This pattern consists of a broad and intense (very bright) central maximum plus a number of narrower and less intense maxima (called **secondary** or **side** maxima) to both sides. In between the maxima are minima. Light flares into those dark regions, but the light waves cancel out one another.

Such a pattern would be totally unexpected in geometrical optics: If light traveled in straight lines as rays, then the slit would allow some of those rays through to form a sharp rendition of the slit on the viewing screen instead of a pattern of bright and dark bands as we see in Fig. 36-1. As in Chapter 35, we must conclude that geometrical optics is only an approximation.

Diffraction is not limited to situations when light passes through a narrow opening (such as a slit or pinhole). It also occurs when light passes an edge, such as the edges of the razor blade whose diffraction pattern is shown in Fig. 36-2. Note the lines of maxima and minima that run approximately parallel to the edges, at both the inside edges of the blade and the outside edges. As the light passes, say, the vertical edge at the left, it flares left and right and undergoes interference, producing the pattern along the left edge. The rightmost portion of that pattern actually lies behind the blade, within what would be the blade's shadow if geometrical optics prevailed.

36-2 DIFFRACTION AND THE WAVE THEORY OF LIGHT

991

You encounter a common example of diffraction when you look at a clear blue sky and see tiny specks and hairlike structures floating in your view. These *floaters*, as they are called, are produced when light passes the edges of tiny deposits in the vitreous humor, the transparent material filling most of the eyeball. What you are seeing when a floater is in your field of vision is the diffraction pattern produced on the retina by one of these deposits. If you sight through a pinhole in a piece of cardboard so as to make the light entering your eye approximately a plane wave, you can distinguish individual maxima and minima in the patterns.

Diffraction is a wave effect. That is, it occurs because light is a wave and it occurs with other types of waves as well. For example, you have probably seen diffraction in action at football games. When a cheerleader near the playing field yells up at several thousand noisy fans, the yell can hardly be heard because the sound waves diffract when they pass through the narrow opening of the cheerleader's mouth. This flaring leaves little of the waves traveling toward the fans in front of the cheerleader. To offset the diffraction, the cheerleader can yell through a megaphone. The sound waves then emerge from the much wider opening at the end of the megaphone. The flaring is thus reduced, and much more of the sound reaches the fans in front of the cheerleader.

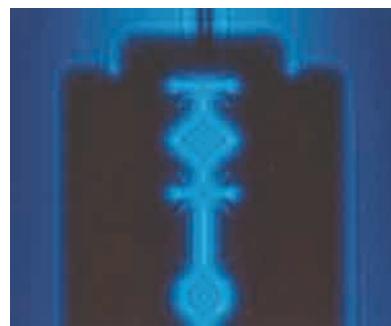


Fig. 36-2 The diffraction pattern produced by a razor blade in monochromatic light. Note the lines of alternating maximum and minimum intensity. (*Ken Kay/Fundamental Photographs*)

The Fresnel Bright Spot

Diffraction finds a ready explanation in the wave theory of light. However, this theory, originally advanced in the late 1600s by Huygens and used 123 years later by Young to explain double-slit interference, was very slow in being adopted, largely because it ran counter to Newton's theory that light was a stream of particles.

Newton's view was the prevailing view in French scientific circles of the early 19th century, when Augustin Fresnel was a young military engineer. Fresnel, who believed in the wave theory of light, submitted a paper to the French Academy of Sciences describing his experiments with light and his wave-theory explanations of them.

In 1819, the Academy, dominated by supporters of Newton and thinking to challenge the wave point of view, organized a prize competition for an essay on the subject of diffraction. Fresnel won. The Newtonians, however, were not swayed. One of them, S. D. Poisson, pointed out the "strange result" that if Fresnel's theories were correct, then light waves should flare into the shadow region of a sphere as they pass the edge of the sphere, producing a bright spot at the center of the shadow. The prize committee arranged a test of Poisson's prediction and discovered that the predicted *Fresnel bright spot*, as we call it today, was indeed there (Fig. 36-3). Nothing builds confidence in a theory so much as having one of its unexpected and counterintuitive predictions verified by experiment.



Fig. 36-3 A photograph of the diffraction pattern of a disk. Note the concentric diffraction rings and the Fresnel bright spot at the center of the pattern. This experiment is essentially identical to that arranged by the committee testing Fresnel's theories, because both the sphere they used and the disk used here have a cross section with a circular edge. (*Jearl Walker*)

36-3 Diffraction by a Single Slit: Locating the Minima

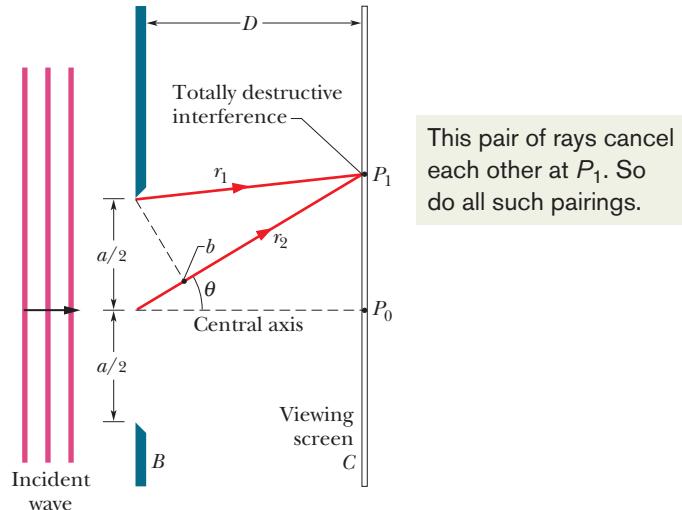
Let us now examine the diffraction pattern of plane waves of light of wavelength λ that are diffracted by a single long, narrow slit of width a in an otherwise opaque screen B , as shown in cross section in Fig. 36-4. (In that figure, the slit's length extends into and out of the page, and the incoming wavefronts are parallel to screen B .) When the diffracted light reaches viewing screen C , waves from different points within the slit undergo interference and produce a diffraction pattern of bright and dark fringes (interference maxima and minima) on the screen. To locate the fringes, we shall use a procedure somewhat similar to the one we used to locate the fringes in a two-slit interference pattern. However, diffraction is more mathematically challenging, and here we shall be able to find equations for only the dark fringes.

Before we do that, however, we can justify the central bright fringe seen in Fig. 36-1 by noting that the Huygens wavelets from all points in the slit travel about the same distance to reach the center of the pattern and thus are in phase there. As for the other bright fringes, we can say only that they are approximately halfway between adjacent dark fringes.

To find the dark fringes, we shall use a clever (and simplifying) strategy that involves pairing up all the rays coming through the slit and then finding what conditions cause the wavelets of the rays in each pair to cancel each other. We apply this strategy in Fig. 36-4 to locate the first dark fringe, at point P_1 . First, we mentally divide the slit into two *zones* of equal widths $a/2$. Then we extend to P_1 a light ray r_1 from the top point of the top zone and a light ray r_2 from the top point of the bottom zone. We want the wavelets along these two rays to cancel each other when they arrive at P_1 . Then any similar pairing of rays from the two zones will give cancellation. A central axis is drawn from the center of the slit to screen C , and P_1 is located at an angle θ to that axis.

The wavelets of the pair of rays r_1 and r_2 are in phase within the slit because they originate from the same wavefront passing through the slit, along the width of the slit. However, to produce the first dark fringe they must be out of phase by $\lambda/2$ when they reach P_1 ; this phase difference is due to their path length difference, with the path traveled by the wavelet of r_2 to reach P_1 being longer than the path traveled by the wavelet of r_1 . To display this path length difference, we find a point b on ray r_2 such that the path length from b to P_1 matches the path length of ray r_1 . Then the path length difference between the two rays is the distance from the center of the slit to b .

Fig. 36-4 Waves from the top points of two zones of width $a/2$ undergo fully destructive interference at point P_1 on viewing screen C .



36-3 DIFFRACTION BY A SINGLE SLIT: LOCATING THE MINIMA

993

When viewing screen C is near screen B , as in Fig. 36-4, the diffraction pattern on C is difficult to describe mathematically. However, we can simplify the mathematics considerably if we arrange for the screen separation D to be much larger than the slit width a . Then we can approximate rays r_1 and r_2 as being parallel, at angle θ to the central axis (Fig. 36-5). We can also approximate the triangle formed by point b , the top point of the slit, and the center point of the slit as being a right triangle, and one of the angles inside that triangle as being θ . The path length difference between rays r_1 and r_2 (which is still the distance from the center of the slit to point b) is then equal to $(a/2) \sin \theta$.

We can repeat this analysis for any other pair of rays originating at corresponding points in the two zones (say, at the midpoints of the zones) and extending to point P_1 . Each such pair of rays has the same path length difference $(a/2) \sin \theta$. Setting this common path length difference equal to $\lambda/2$ (our condition for the first dark fringe), we have

$$\frac{a}{2} \sin \theta = \frac{\lambda}{2},$$

which gives us

$$a \sin \theta = \lambda \quad (\text{first minimum}). \quad (36-1)$$

Given slit width a and wavelength λ , Eq. 36-1 tells us the angle θ of the first dark fringe above and (by symmetry) below the central axis.

Note that if we begin with $a > \lambda$ and then narrow the slit while holding the wavelength constant, we increase the angle at which the first dark fringes appear; that is, the extent of the diffraction (the extent of the flaring and the width of the pattern) is *greater* for a *narrower* slit. When we have reduced the slit width to the wavelength (that is, $a = \lambda$), the angle of the first dark fringes is 90° . Since the first dark fringes mark the two edges of the central bright fringe, that bright fringe must then cover the entire viewing screen.

We find the second dark fringes above and below the central axis as we found the first dark fringes, except that we now divide the slit into *four* zones of equal widths $a/4$, as shown in Fig. 36-6a. We then extend rays r_1 , r_2 , r_3 , and r_4 from the top points of the zones to point P_2 , the location of the second dark fringe above the central axis. To produce that fringe, the path length difference

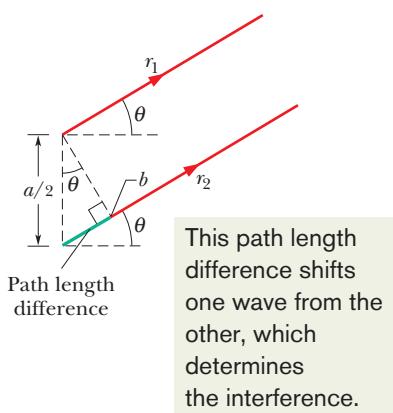


Fig. 36-5 For $D \gg a$, we can approximate rays r_1 and r_2 as being parallel, at angle θ to the central axis.

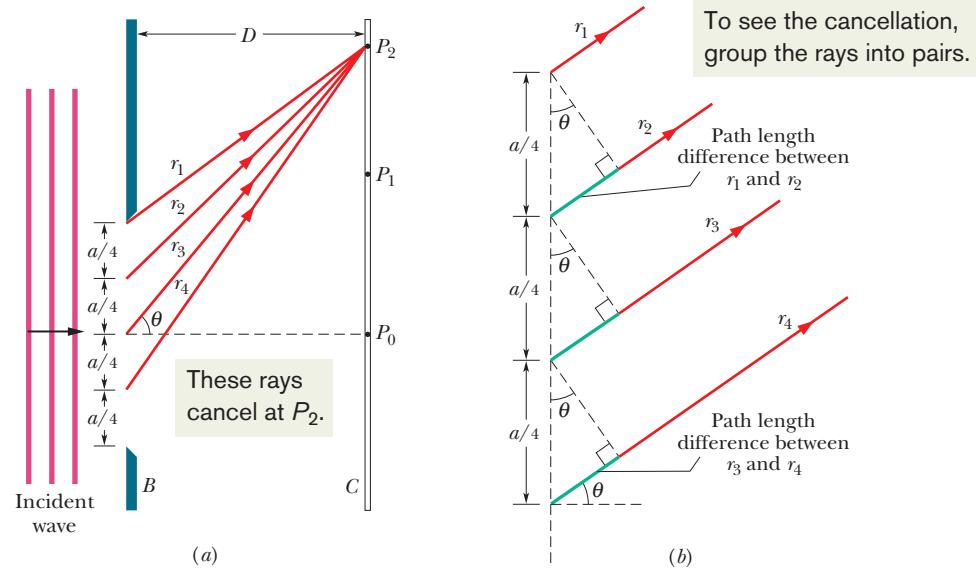


Fig. 36-6 (a) Waves from the top points of four zones of width $a/4$ undergo fully destructive interference at point P_2 . (b) For $D \gg a$, we can approximate rays r_1, r_2, r_3 , and r_4 as being parallel, at angle θ to the central axis.

between r_1 and r_2 , that between r_2 and r_3 , and that between r_3 and r_4 must all be equal to $\lambda/2$.

For $D \gg a$, we can approximate these four rays as being parallel, at angle θ to the central axis. To display their path length differences, we extend a perpendicular line through each adjacent pair of rays, as shown in Fig. 36-6b, to form a series of right triangles, each of which has a path length difference as one side. We see from the top triangle that the path length difference between r_1 and r_2 is $(a/4) \sin \theta$. Similarly, from the bottom triangle, the path length difference between r_3 and r_4 is also $(a/4) \sin \theta$. In fact, the path length difference for any two rays that originate at corresponding points in two adjacent zones is $(a/4) \sin \theta$. Since in each such case the path length difference is equal to $\lambda/2$, we have

$$\frac{a}{4} \sin \theta = \frac{\lambda}{2},$$

which gives us

$$a \sin \theta = 2\lambda \quad (\text{second minimum}). \quad (36-2)$$

We could now continue to locate dark fringes in the diffraction pattern by splitting up the slit into more zones of equal width. We would always choose an even number of zones so that the zones (and their waves) could be paired as we have been doing. We would find that the dark fringes above and below the central axis can be located with the general equation

$$a \sin \theta = m\lambda, \quad \text{for } m = 1, 2, 3, \dots \quad (\text{minima—dark fringes}). \quad (36-3)$$

You can remember this result in the following way. Draw a triangle like the one in Fig. 36-5, but for the full slit width a , and note that the path length difference between the top and bottom rays equals $a \sin \theta$. Thus, Eq. 36-3 says:



In a single-slit diffraction experiment, dark fringes are produced where the path length differences ($a \sin \theta$) between the top and bottom rays are equal to $\lambda, 2\lambda, 3\lambda, \dots$

This may seem to be wrong because the waves of those two particular rays will be exactly in phase with each other when their path length difference is an integer number of wavelengths. However, they each will still be part of a pair of waves that are exactly out of phase with each other; thus, *each* wave will be canceled by some other wave, resulting in darkness. (Two light waves that are exactly out of phase will always cancel each other, giving a net wave of zero, even if they happen to be exactly in phase with other light waves.)

Equations 36-1, 36-2, and 36-3 are derived for the case of $D \gg a$. However, they also apply if we place a converging lens between the slit and the viewing screen and then move the screen in so that it coincides with the focal plane of the lens. The lens ensures that rays which now reach any point on the screen are *exactly* parallel (rather than approximately) back at the slit. They are like the initially parallel rays of Fig. 34-14a that are directed to the focal point by a converging lens.



CHECKPOINT 1

We produce a diffraction pattern on a viewing screen by means of a long narrow slit illuminated by blue light. Does the pattern expand away from the bright center (the maxima and minima shift away from the center) or contract toward it if we (a) switch to yellow light or (b) decrease the slit width?

Sample Problem**Single-slit diffraction pattern with white light**

A slit of width a is illuminated by white light.

- (a) For what value of a will the first minimum for red light of wavelength $\lambda = 650 \text{ nm}$ appear at $\theta = 15^\circ$?

KEY IDEA

Diffraction occurs separately for each wavelength in the range of wavelengths passing through the slit, with the locations of the minima for each wavelength given by Eq. 36-3 ($a \sin \theta = m\lambda$).

Calculation: When we set $m = 1$ (for the first minimum) and substitute the given values of θ and λ , Eq. 36-3 yields

$$\begin{aligned} a &= \frac{m\lambda}{\sin \theta} = \frac{(1)(650 \text{ nm})}{\sin 15^\circ} \\ &= 2511 \text{ nm} \approx 2.5 \mu\text{m}. \end{aligned} \quad (\text{Answer})$$

For the incident light to flare out that much ($\pm 15^\circ$ to the first minima) the slit has to be very fine indeed—in this case, a mere four times the wavelength. For comparison, note that a fine human hair may be about $100 \mu\text{m}$ in diameter.

- (b) What is the wavelength λ' of the light whose first side diffraction maximum is at 15° , thus coinciding with the first minimum for the red light?

KEY IDEA

The first side maximum for any wavelength is about halfway between the first and second minima for that wavelength.

Calculations: Those first and second minima can be located with Eq. 36-3 by setting $m = 1$ and $m = 2$, respectively. Thus, the first side maximum can be located approximately by setting $m = 1.5$. Then Eq. 36-3 becomes

$$a \sin \theta = 1.5\lambda'.$$

Solving for λ' and substituting known data yield

$$\begin{aligned} \lambda' &= \frac{a \sin \theta}{1.5} = \frac{(2511 \text{ nm})(\sin 15^\circ)}{1.5} \\ &= 430 \text{ nm}. \end{aligned} \quad (\text{Answer})$$

Light of this wavelength is violet (far blue, near the short-wavelength limit of the human range of visible light). From the two equations we used, can you see that the first side maximum for light of wavelength 430 nm will always coincide with the first minimum for light of wavelength 650 nm, no matter what the slit width is? However, the angle θ at which this overlap occurs does depend on slit width. If the slit is relatively narrow, the angle will be relatively large, and conversely.



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36-4 Intensity in Single-Slit Diffraction, Qualitatively

In Section 36-3 we saw how to find the positions of the minima and the maxima in a single-slit diffraction pattern. Now we turn to a more general problem: find an expression for the intensity I of the pattern as a function of θ , the angular position of a point on a viewing screen.

To do this, we divide the slit of Fig. 36-4 into N zones of equal widths Δx small enough that we can assume each zone acts as a source of Huygens wavelets. We wish to superimpose the wavelets arriving at an arbitrary point P on the viewing screen, at angle θ to the central axis, so that we can determine the amplitude E_θ of the electric component of the resultant wave at P . The intensity of the light at P is then proportional to the square of that amplitude.

To find E_θ , we need the phase relationships among the arriving wavelets. The phase difference between wavelets from adjacent zones is given by

$$\left(\frac{\text{phase}}{\text{difference}} \right) = \left(\frac{2\pi}{\lambda} \right) \left(\frac{\text{path length}}{\text{difference}} \right).$$

For point P at angle θ , the path length difference between wavelets from adjacent zones is $\Delta x \sin \theta$; so the phase difference $\Delta\phi$ between wavelets from adjacent zones is

$$\Delta\phi = \left(\frac{2\pi}{\lambda} \right) (\Delta x \sin \theta). \quad (36-4)$$

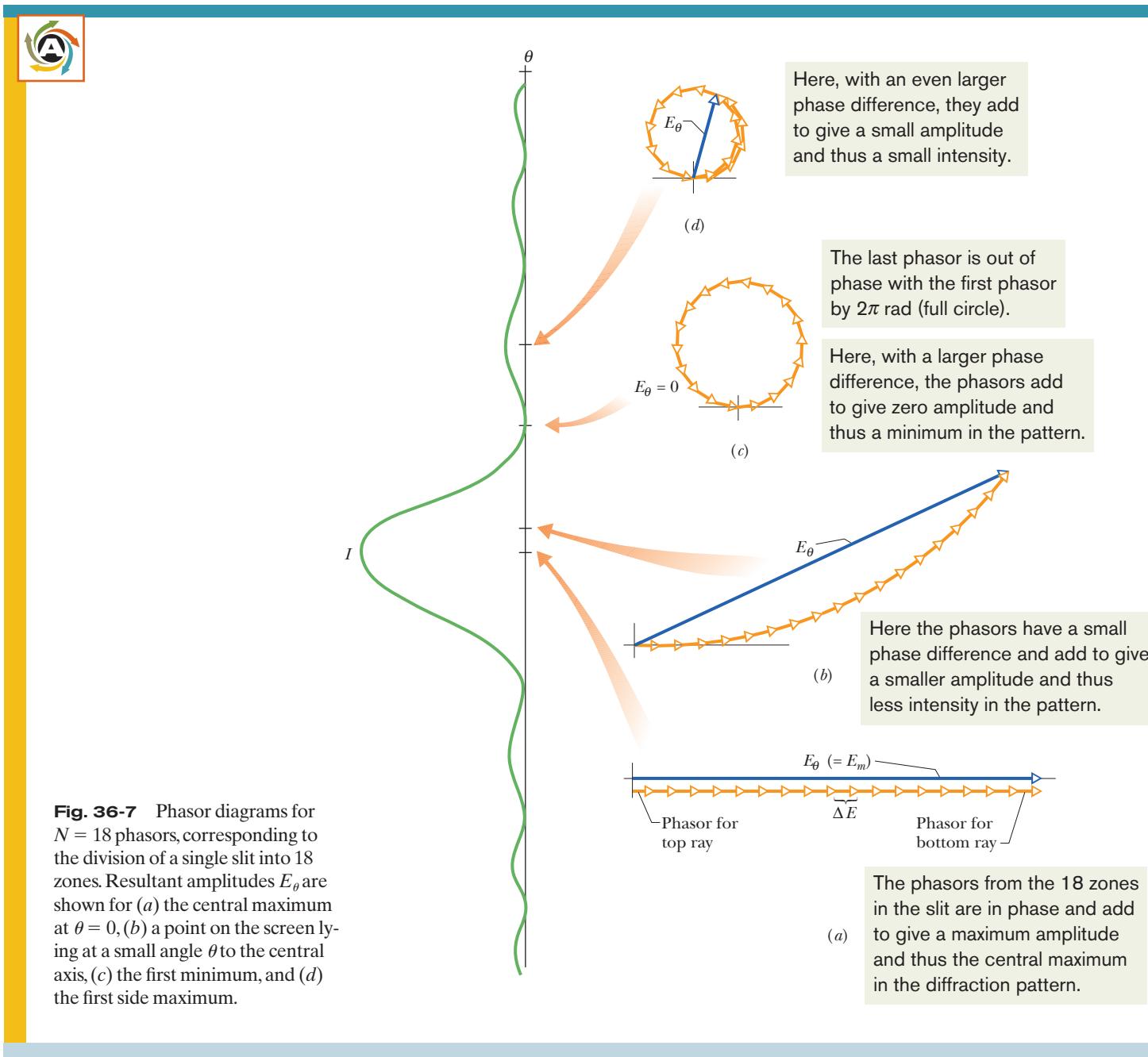


Fig. 36-7 Phasor diagrams for $N = 18$ phasors, corresponding to the division of a single slit into 18 zones. Resultant amplitudes E_θ are shown for (a) the central maximum at $\theta = 0$, (b) a point on the screen lying at a small angle θ to the central axis, (c) the first minimum, and (d) the first side maximum.

We assume that the wavelets arriving at P all have the same amplitude ΔE . To find the amplitude E_θ of the resultant wave at P , we add the amplitude ΔE via phasors. To do this, we construct a diagram of N phasors, one corresponding to the wavelet from each zone in the slit.

For point P_0 at $\theta = 0$ on the central axis of Fig. 36-4, Eq. 36-4 tells us that the phase difference $\Delta\phi$ between the wavelets is zero; that is, the wavelets all arrive in phase. Figure 36-7a is the corresponding phasor diagram; adjacent phasors represent wavelets from adjacent zones and are arranged head to tail. Because there is zero phase difference between the wavelets, there is zero angle between each pair of adjacent phasors. The amplitude E_θ of the net wave at P_0 is the vector sum of these phasors. This arrangement of the phasors turns out to be the one that gives the greatest value for the amplitude E_θ . We call this value E_m ; that is, E_m is the value of E_θ for $\theta = 0$.

36-5 INTENSITY IN SINGLE-SLIT DIFFRACTION, QUANTITATIVELY

997

We next consider a point P that is at a small angle θ to the central axis. Equation 36-4 now tells us that the phase difference $\Delta\phi$ between wavelets from adjacent zones is no longer zero. Figure 36-7b shows the corresponding phasor diagram; as before, the phasors are arranged head to tail, but now there is an angle $\Delta\phi$ between adjacent phasors. The amplitude E_θ at this new point is still the vector sum of the phasors, but it is smaller than that in Fig. 36-7a, which means that the intensity of the light is less at this new point P than at P_0 .

If we continue to increase θ , the angle $\Delta\phi$ between adjacent phasors increases, and eventually the chain of phasors curls completely around so that the head of the last phasor just reaches the tail of the first phasor (Fig. 36-7c). The amplitude E_θ is now zero, which means that the intensity of the light is also zero. We have reached the first minimum, or dark fringe, in the diffraction pattern. The first and last phasors now have a phase difference of 2π rad, which means that the path length difference between the top and bottom rays through the slit equals one wavelength. Recall that this is the condition we determined for the first diffraction minimum.

As we continue to increase θ , the angle $\Delta\phi$ between adjacent phasors continues to increase, the chain of phasors begins to wrap back on itself, and the resulting coil begins to shrink. Amplitude E_θ now increases until it reaches a maximum value in the arrangement shown in Fig. 36-7d. This arrangement corresponds to the first side maximum in the diffraction pattern.

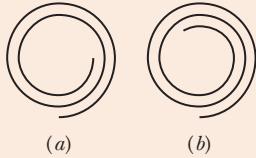
If we increase θ a bit more, the resulting shrinkage of the coil decreases E_θ , which means that the intensity also decreases. When θ is increased enough, the head of the last phasor again meets the tail of the first phasor. We have then reached the second minimum.

We could continue this qualitative method of determining the maxima and minima of the diffraction pattern but, instead, we shall now turn to a quantitative method.



CHECKPOINT 2

The figures represent, in smoother form (with more phasors) than Fig. 36-7, the phasor diagrams for two points of a diffraction pattern that are on opposite sides of a certain diffraction maximum. (a) Which maximum is it? (b) What is the approximate value of m (in Eq. 36-3) that corresponds to this maximum?



36-5 Intensity in Single-Slit Diffraction, Quantitatively

Equation 36-3 tells us how to locate the minima of the single-slit diffraction pattern on screen C of Fig. 36-4 as a function of the angle θ in that figure. Here we wish to derive an expression for the intensity $I(\theta)$ of the pattern as a function of θ . We state, and shall prove below, that the intensity is given by

$$I(\theta) = I_m \left(\frac{\sin \alpha}{\alpha} \right)^2, \quad (36-5)$$

where

$$\alpha = \frac{1}{2}\phi = \frac{\pi a}{\lambda} \sin \theta. \quad (36-6)$$

The symbol α is just a convenient connection between the angle θ that locates a point on the viewing screen and the light intensity $I(\theta)$ at that point. The intensity I_m is the greatest value of the intensities $I(\theta)$ in the pattern and occurs at the central maximum (where $\theta = 0$), and ϕ is the phase difference (in radians) between the top and bottom rays from the slit of width a .

Study of Eq. 36-5 shows that intensity minima will occur where

$$\alpha = m\pi, \quad \text{for } m = 1, 2, 3, \dots \quad (36-7)$$

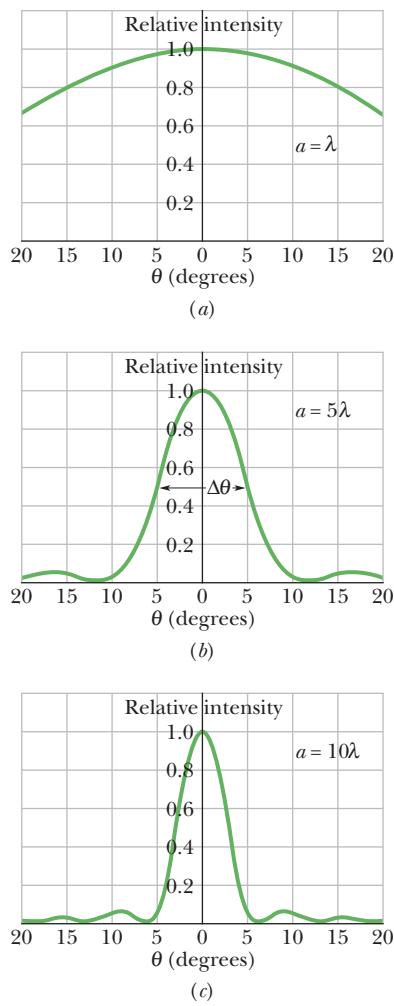


Fig. 36-8 The relative intensity in single-slit diffraction for three values of the ratio a/λ . The wider the slit is, the narrower is the central diffraction maximum.

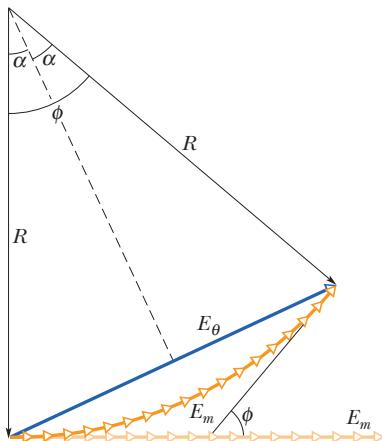


Fig. 36-9 A construction used to calculate the intensity in single-slit diffraction. The situation shown corresponds to that of Fig. 36-7b.

If we put this result into Eq. 36-6, we find

$$m\pi = \frac{\pi a}{\lambda} \sin \theta, \quad \text{for } m = 1, 2, 3, \dots,$$

$$\text{or} \quad a \sin \theta = m\lambda, \quad \text{for } m = 1, 2, 3, \dots \quad (\text{minima—dark fringes}), \quad (36-8)$$

which is exactly Eq. 36-3, the expression that we derived earlier for the location of the minima.

Figure 36-8 shows plots of the intensity of a single-slit diffraction pattern, calculated with Eqs. 36-5 and 36-6 for three slit widths: $a = \lambda$, $a = 5\lambda$, and $a = 10\lambda$. Note that as the slit width increases (relative to the wavelength), the width of the *central diffraction maximum* (the central hill-like region of the graphs) decreases; that is, the light undergoes less flaring by the slit. The secondary maxima also decrease in width (and become weaker). In the limit of slit width a being much greater than wavelength λ , the secondary maxima due to the slit disappear; we then no longer have single-slit diffraction (but we still have diffraction due to the edges of the wide slit, like that produced by the edges of the razor blade in Fig. 36-2).

Proof of Eqs. 36-5 and 36-6

To find an expression for the intensity at a point in the diffraction pattern, we need to divide the slit into many zones and then add the phasors corresponding to those zones, as we did in Fig. 36-7. The arc of phasors in Fig. 36-9 represents the wavelets that reach an arbitrary point P on the viewing screen of Fig. 36-4, corresponding to a particular small angle θ . The amplitude E_θ of the resultant wave at P is the vector sum of these phasors. If we divide the slit of Fig. 36-4 into infinitesimal zones of width Δx , the arc of phasors in Fig. 36-9 approaches the arc of a circle; we call its radius R as indicated in that figure. The length of the arc must be E_m , the amplitude at the center of the diffraction pattern, because if we straightened out the arc we would have the phasor arrangement of Fig. 36-7a (shown lightly in Fig. 36-9).

The angle ϕ in the lower part of Fig. 36-9 is the difference in phase between the infinitesimal vectors at the left and right ends of arc E_m . From the geometry, ϕ is also the angle between the two radii marked R in Fig. 36-9. The dashed line in that figure, which bisects ϕ , then forms two congruent right triangles. From either triangle we can write

$$\sin \frac{1}{2}\phi = \frac{E_\theta}{2R}. \quad (36-9)$$

In radian measure, ϕ is (with E_m considered to be a circular arc)

$$\phi = \frac{E_m}{R}.$$

Solving this equation for R and substituting in Eq. 36-9 lead to

$$E_\theta = \frac{E_m}{\frac{1}{2}\phi} \sin \frac{1}{2}\phi. \quad (36-10)$$

In Section 33-5 we saw that the intensity of an electromagnetic wave is proportional to the square of the amplitude of its electric field. Here, this means that the maximum intensity I_m (which occurs at the center of the diffraction pattern) is proportional to E_m^2 and the intensity $I(\theta)$ at angle θ is proportional to E_θ^2 . Thus, we may write

$$\frac{I(\theta)}{I_m} = \frac{E_\theta^2}{E_m^2}. \quad (36-11)$$

Substituting for E_θ with Eq. 36-10 and then substituting $\alpha = \frac{1}{2}\phi$, we are led to the

36-5 INTENSITY IN SINGLE-SLIT DIFFRACTION, QUANTITATIVELY

999

following expression for the intensity as a function of θ :

$$I(\theta) = I_m \left(\frac{\sin \alpha}{\alpha} \right)^2.$$

This is exactly Eq. 36-5, one of the two equations we set out to prove.

The second equation we wish to prove relates α to θ . The phase difference ϕ between the rays from the top and bottom of the entire slit may be related to a path length difference with Eq. 36-4; it tells us that

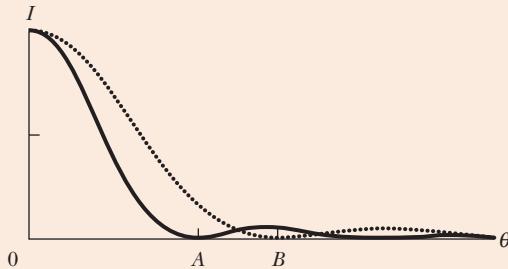
$$\phi = \left(\frac{2\pi}{\lambda} \right) (a \sin \theta),$$

where a is the sum of the widths Δx of the infinitesimal zones. However, $\phi = 2\alpha$, so this equation reduces to Eq. 36-6.



CHECKPOINT 3

Two wavelengths, 650 and 430 nm, are used separately in a single-slit diffraction experiment. The figure shows the results as graphs of intensity I versus angle θ for the two diffraction patterns. If both wavelengths are then used simultaneously, what color will be seen in the combined diffraction pattern at (a) angle A and (b) angle B' ?



Sample Problem

Intensities of the maxima in a single-slit interference pattern

Find the intensities of the first three secondary maxima (side maxima) in the single-slit diffraction pattern of Fig. 36-1, measured as a percentage of the intensity of the central maximum.

KEY IDEAS

The secondary maxima lie approximately halfway between the minima, whose angular locations are given by Eq. 36-7 ($\alpha = m\pi$). The locations of the secondary maxima are then given (approximately) by

$$a = (m + \frac{1}{2})\pi, \quad \text{for } m = 1, 2, 3, \dots,$$

with α in radian measure. We can relate the intensity I at any point in the diffraction pattern to the intensity I_m of the central maximum via Eq. 36-5.

Calculations: Substituting the approximate values of α for the secondary maxima into Eq. 36-5 to obtain the relative

intensities at those maxima, we get

$$\frac{I}{I_m} = \left(\frac{\sin \alpha}{\alpha} \right)^2 = \left(\frac{\sin(m + \frac{1}{2})\pi}{(m + \frac{1}{2})\pi} \right)^2, \quad \text{for } m = 1, 2, 3, \dots$$

The first of the secondary maxima occurs for $m = 1$, and its relative intensity is

$$\begin{aligned} \frac{I_1}{I_m} &= \left(\frac{\sin(1 + \frac{1}{2})\pi}{(1 + \frac{1}{2})\pi} \right)^2 = \left(\frac{\sin 1.5\pi}{1.5\pi} \right)^2 \\ &= 4.50 \times 10^{-2} \approx 4.5\%. \end{aligned} \quad (\text{Answer})$$

For $m = 2$ and $m = 3$ we find that

$$\frac{I_2}{I_m} = 1.6\% \quad \text{and} \quad \frac{I_3}{I_m} = 0.83\%. \quad (\text{Answer})$$

As you can see from these results, successive secondary maxima decrease rapidly in intensity. Figure 36-1 was deliberately overexposed to reveal them.



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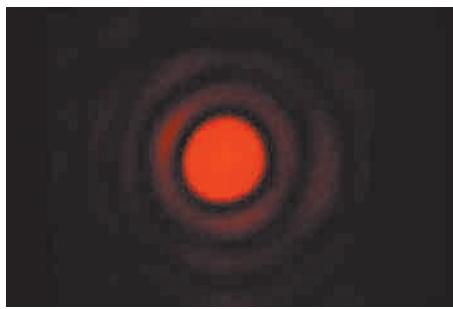


Fig. 36-10 The diffraction pattern of a circular aperture. Note the central maximum and the circular secondary maxima. The figure has been overexposed to bring out these secondary maxima, which are much less intense than the central maximum. (*Jearl Walker*)

36-6 Diffraction by a Circular Aperture

Here we consider diffraction by a circular aperture—that is, a circular opening, such as a circular lens, through which light can pass. Figure 36-10 shows the image formed by light from a laser that was directed onto a circular aperture with a very small diameter. This image is not a point, as geometrical optics would suggest, but a circular disk surrounded by several progressively fainter secondary rings. Comparison with Fig. 36-1 leaves little doubt that we are dealing with a diffraction phenomenon. Here, however, the aperture is a circle of diameter d rather than a rectangular slit.

The (complex) analysis of such patterns shows that the first minimum for the diffraction pattern of a circular aperture of diameter d is located by

$$\sin \theta = 1.22 \frac{\lambda}{d} \quad (\text{first minimum—circular aperture}). \quad (36-12)$$

The angle θ here is the angle from the central axis to any point on that (circular) minimum. Compare this with Eq. 36-1,

$$\sin \theta = \frac{\lambda}{a} \quad (\text{first minimum—single slit}), \quad (36-13)$$

which locates the first minimum for a long narrow slit of width a . The main difference is the factor 1.22, which enters because of the circular shape of the aperture.

Resolvability

The fact that lens images are diffraction patterns is important when we wish to *resolve* (distinguish) two distant point objects whose angular separation is small. Figure 36-11 shows, in three different cases, the visual appearance and corresponding intensity pattern for two distant point objects (stars, say) with small angular separation. In Figure 36-11a, the objects are not resolved because of diffraction; that is, their diffraction patterns (mainly their central maxima) overlap so much that the two objects cannot be distinguished from a single point object. In Fig. 36-11b the objects are barely resolved, and in Fig. 36-11c they are fully resolved.

In Fig. 36-11b the angular separation of the two point sources is such that the central maximum of the diffraction pattern of one source is centered on the first minimum of the diffraction pattern of the other, a condition called **Rayleigh's criterion** for resolvability. From Eq. 36-12, two objects that are barely resolvable

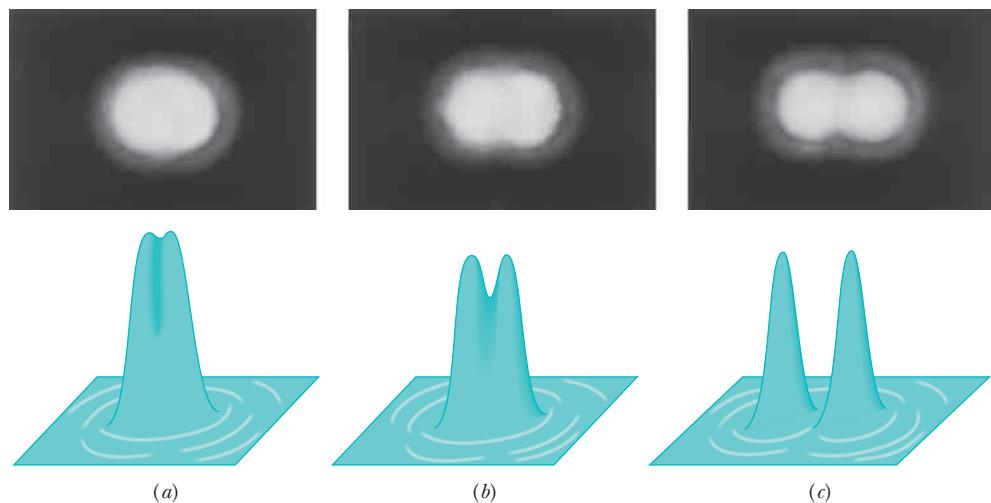


Fig. 36-11 At the top, the images of two point sources (stars) formed by a converging lens. At the bottom, representations of the image intensities. In (a) the angular separation of the sources is too small for them to be distinguished, in (b) they can be marginally distinguished, and in (c) they are clearly distinguished. Rayleigh's criterion is satisfied in (b), with the central maximum of one diffraction pattern coinciding with the first minimum of the other.

36-6 DIFFRACTION BY A CIRCULAR APERTURE

1001

by this criterion must have an angular separation θ_R of

$$\theta_R = \sin^{-1} \frac{1.22\lambda}{d}.$$

Since the angles are small, we can replace $\sin \theta_R$ with θ_R expressed in radians:

$$\theta_R = 1.22 \frac{\lambda}{d} \quad (\text{Rayleigh's criterion}). \quad (36-14)$$

Applying Rayleigh's criterion for resolvability to human vision is only an approximation because visual resolvability depends on many factors, such as the relative brightness of the sources and their surroundings, turbulence in the air between the sources and the observer, and the functioning of the observer's visual system. Experimental results show that the least angular separation that can actually be resolved by a person is generally somewhat greater than the value given by Eq. 36-14. However, for calculations here, we shall take Eq. 36-14 as being a precise criterion: If the angular separation θ between the sources is greater than θ_R , we can visually resolve the sources; if it is less, we cannot.

Rayleigh's criterion can explain the arresting illusions of color in the style of painting known as pointillism (Fig. 36-12). In this style, a painting is made not with brush strokes in the usual sense but rather with a myriad of small colored dots. One fascinating aspect of a pointillistic painting is that when you change your distance from it, the colors shift in subtle, almost subconscious ways. This color shifting has to do with whether you can resolve the colored dots. When you stand close enough to the painting, the angular separations θ of adjacent dots are greater than θ_R and thus the dots can be seen individually. Their colors are the true colors of the paints used. However, when you stand far enough from the painting, the angular separations θ are less than θ_R and the dots cannot be seen individually. The resulting blend of colors coming into your eye from any group of dots can then cause your brain to "make up" a color for that group—a color that may not actually exist in the group. In this way, a pointillistic painter uses your visual system to create the colors of the art.

When we wish to use a lens instead of our visual system to resolve objects of small angular separation, it is desirable to make the diffraction pattern as small as possible. According to Eq. 36-14, this can be done either by increasing the lens diameter or by using light of a shorter wavelength. For this reason ultraviolet light is often used with microscopes because its wavelength is shorter than a visible light wavelength.



Fig. 36-12 The pointillistic painting *The Seine at Herblay* by Maximilien Luce consists of thousands of colored dots. With the viewer very close to the canvas, the dots and their true colors are visible. At normal viewing distances, the dots are irresolvable and thus blend. (*Maximilien Luce, The Seine at Herblay, 1890. Musee d'Orsay, Paris, France. Photo by Erich Lessing/Art Resource*)

CHECKPOINT 4

Suppose that you can barely resolve two red dots because of diffraction by the pupil of your eye. If we increase the general illumination around you so that the pupil decreases in diameter, does the resolvability of the dots improve or diminish? Consider only diffraction. (You might experiment to check your answer.)

Sample Problem**Pointillistic paintings use the diffraction of your eye**

Figure 36-13a is a representation of the colored dots on a pointillistic painting. Assume that the average center-to-center separation of the dots is $D = 2.0 \text{ mm}$. Also assume that the diameter of the pupil of your eye is $d = 1.5 \text{ mm}$ and that the least angular separation between dots you can resolve is set only by Rayleigh's criterion. What is the least viewing distance from which you cannot distinguish any dots on the painting?

**KEY IDEA**

Consider any two adjacent dots that you can distinguish when you are close to the painting. As you move away, you continue to distinguish the dots until their angular separation θ (in your view) has decreased to the angle given by

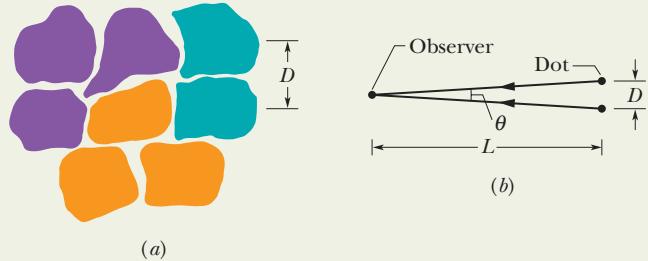


Fig. 36-13 (a) Representation of some dots on a pointillistic painting, showing an average center-to-center separation D . (b) The arrangement of separation D between two dots, their angular separation θ , and the viewing distance L .

Rayleigh's criterion:

$$\theta_R = 1.22 \frac{\lambda}{d}. \quad (36-15)$$

Calculations: Figure 36-13b shows, from the side, the angular separation θ of the dots, their center-to-center separation D , and your distance L from them. Because D/L is small, angle θ is also small and we can make the approximation

$$\theta = \frac{D}{L}. \quad (36-16)$$

Setting θ of Eq. 36-16 equal to θ_R of Eq. 36-15 and solving for L , we then have

$$L = \frac{Dd}{1.22\lambda}. \quad (36-17)$$

Equation 36-17 tells us that L is larger for smaller λ . Thus, as you move away from the painting, adjacent red dots (long wavelengths) become indistinguishable before adjacent blue dots do. To find the least distance L at which *no* colored dots are distinguishable, we substitute $\lambda = 400 \text{ nm}$ (blue or violet light) into Eq. 36-17:

$$L = \frac{(2.0 \times 10^{-3} \text{ m})(1.5 \times 10^{-3} \text{ m})}{(1.22)(400 \times 10^{-9} \text{ m})} = 6.1 \text{ m. (Answer)}$$

At this or a greater distance, the color you perceive at any given spot on the painting is a blended color that may not actually exist there.

Sample Problem**Rayleigh's criterion for resolving two distant objects**

A circular converging lens, with diameter $d = 32 \text{ mm}$ and focal length $f = 24 \text{ cm}$, forms images of distant point objects in the focal plane of the lens. The wavelength is $\lambda = 550 \text{ nm}$.

- (a) Considering diffraction by the lens, what angular separation must two distant point objects have to satisfy Rayleigh's criterion?

KEY IDEA

Figure 36-14 shows two distant point objects P_1 and P_2 , the lens, and a viewing screen in the focal plane of the lens. It also shows, on the right, plots of light intensity I versus position on the screen for the central maxima of the images formed by the lens. Note that the angular separation θ_o of the objects equals the angular separation θ_i of the images. Thus, if the images are to satisfy Rayleigh's criterion

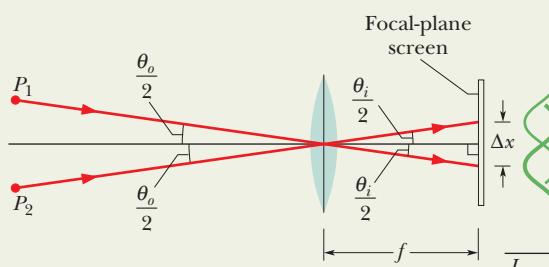


Fig. 36-14 Light from two distant point objects P_1 and P_2 passes through a converging lens and forms images on a viewing screen in the focal plane of the lens. Only one representative ray from each object is shown. The images are not points but diffraction patterns, with intensities approximately as plotted at the right. The angular separation of the objects is θ_o and that of the images is θ_i ; the central maxima of the images have a separation Δx .

for resolvability, the angular separations on both sides of the lens must be given by Eq. 36-14 (assuming small angles).

Calculations: From Eq. 36-14, we obtain

$$\theta_o = \theta_i = \theta_R = 1.22 \frac{\lambda}{d}$$

$$= \frac{(1.22)(550 \times 10^{-9} \text{ m})}{32 \times 10^{-3} \text{ m}} = 2.1 \times 10^{-5} \text{ rad. (Answer)}$$

At this angular separation, each central maximum in the two intensity curves of Fig. 36-14 is centered on the first minimum of the other curve.



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- (b) What is the separation Δx of the centers of the *images* in the focal plane? (That is, what is the separation of the *central* peaks in the two intensity-versus-position curves?)

Calculations: From either triangle between the lens and the screen in Fig. 36-14, we see that $\tan \theta_i/2 = \Delta x/2f$. Rearranging this equation and making the approximation $\tan \theta \approx \theta$, we find

$$\Delta x = f\theta_i, \quad (36-18)$$

where θ_i is in radian measure. Substituting known data then yields

$$\Delta x = (0.24 \text{ m})(2.1 \times 10^{-5} \text{ rad}) = 5.0 \mu\text{m.} \quad (\text{Answer})$$

36-7 Diffraction by a Double Slit

In the double-slit experiments of Chapter 35, we implicitly assumed that the slits were much narrower than the wavelength of the light illuminating them; that is, $a \ll \lambda$. For such narrow slits, the central maximum of the diffraction pattern of either slit covers the entire viewing screen. Moreover, the interference of light from the two slits produces bright fringes with approximately the same intensity (Fig. 35-12).

In practice with visible light, however, the condition $a \ll \lambda$ is often not met. For relatively wide slits, the interference of light from two slits produces bright fringes that do not all have the same intensity. That is, the intensities of the fringes produced by double-slit interference (as discussed in Chapter 35) are modified by diffraction of the light passing through each slit (as discussed in this chapter).

As an example, the intensity plot of Fig. 36-15a suggests the double-slit interference pattern that would occur if the slits were infinitely narrow (and thus $a \ll \lambda$); all the bright interference fringes would have the same intensity. The intensity plot of Fig. 36-15b is that for diffraction by a single actual slit; the diffraction pattern has a broad central maximum and weaker secondary maxima at $\pm 17^\circ$. The plot of Fig. 36-15c suggests the interference pattern for two actual slits. That plot was constructed by using the curve of Fig. 36-15b as an *envelope* on the intensity plot in Fig. 36-15a. The positions of the fringes are not changed; only the intensities are affected.

Figure 36-16a shows an actual pattern in which both double-slit interference and diffraction are evident. If one slit is covered, the single-slit diffraction pattern of Fig. 36-16b results. Note the correspondence between Figs. 36-16a and 36-15c, and between Figs. 36-16b and 36-15b. In comparing these figures, bear in mind that Fig. 36-16 has been deliberately overexposed to bring

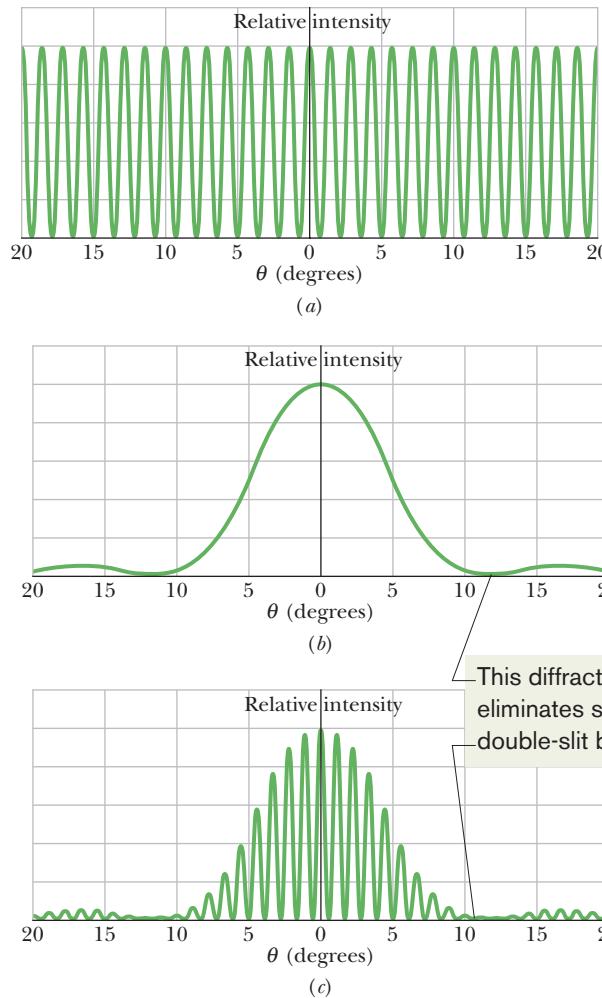


Fig. 36-15 (a) The intensity plot to be expected in a double-slit interference experiment with vanishingly narrow slits. (b) The intensity plot for diffraction by a typical slit of width a (not vanishingly narrow). (c) The intensity plot to be expected for two slits of width a . The curve of (b) acts as an envelope, limiting the intensity of the double-slit fringes in (a). Note that the first minima of the diffraction pattern of (b) eliminate the double-slit fringes that would occur near 12° in (c).

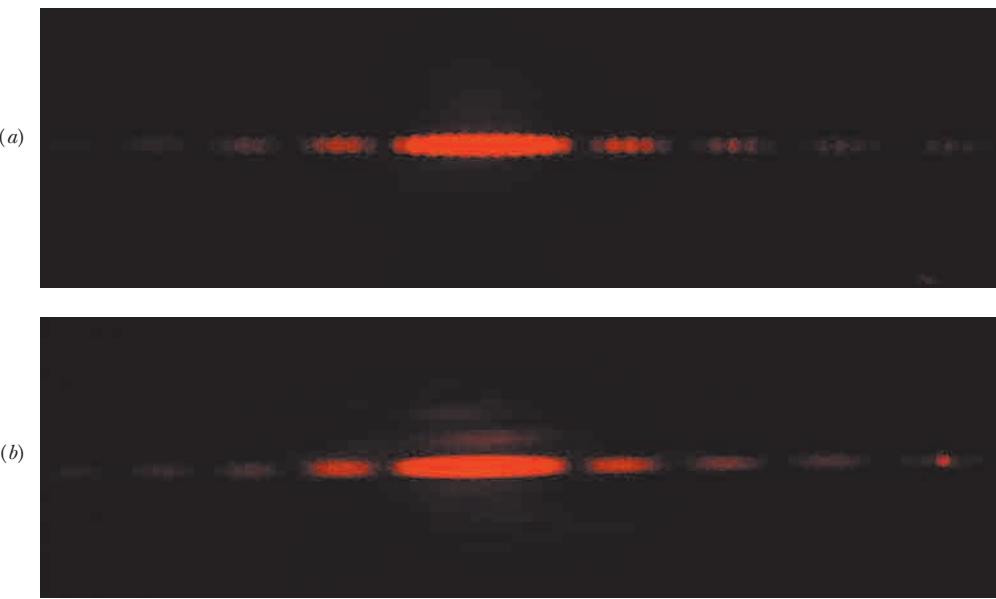


Fig. 36-16 (a) Interference fringes for an actual double-slit system; compare with Fig. 36-15c. (b) The diffraction pattern of a single slit; compare with Fig. 36-15b. (Jearl Walker)

out the faint secondary maxima and that several secondary maxima (rather than one) are shown.

With diffraction effects taken into account, the intensity of a double-slit interference pattern is given by

$$I(\theta) = I_m (\cos^2 \beta) \left(\frac{\sin \alpha}{\alpha} \right)^2 \quad (\text{double slit}), \quad (36-19)$$

in which $\beta = \frac{\pi d}{\lambda} \sin \theta$ (36-20)

and $\alpha = \frac{\pi a}{\lambda} \sin \theta$. (36-21)

Here d is the distance between the centers of the slits and a is the slit width. Note carefully that the right side of Eq. 36-19 is the product of I_m and two factors. (1) The *interference factor* $\cos^2 \beta$ is due to the interference between two slits with slit separation d (as given by Eqs. 35-22 and 35-23). (2) The *diffraction factor* $[(\sin \alpha)/\alpha]^2$ is due to diffraction by a single slit of width a (as given by Eqs. 36-5 and 36-6).

Let us check these factors. If we let $a \rightarrow 0$ in Eq. 36-21, for example, then $\alpha \rightarrow 0$ and $(\sin \alpha)/\alpha \rightarrow 1$. Equation 36-19 then reduces, as it must, to an equation describing the interference pattern for a pair of vanishingly narrow slits with slit separation d . Similarly, putting $d = 0$ in Eq. 36-20 is equivalent physically to causing the two slits to merge into a single slit of width a . Then Eq. 36-20 yields $\beta = 0$ and $\cos^2 \beta = 1$. In this case Eq. 36-19 reduces, as it must, to an equation describing the diffraction pattern for a single slit of width a .

The double-slit pattern described by Eq. 36-19 and displayed in Fig. 36-16a combines interference and diffraction in an intimate way. Both are superposition effects, in that they result from the combining of waves with different phases at a given point. If the combining waves originate from a small number of elementary coherent sources—as in a double-slit experiment with $a \ll \lambda$ —we call the process *interference*. If the combining waves originate in a single wavefront—as in

a single-slit experiment—we call the process *diffraction*. This distinction between interference and diffraction (which is somewhat arbitrary and not always adhered to) is a convenient one, but we should not forget that both are superposition effects and usually both are present simultaneously (as in Fig. 36-16a).

Sample Problem

Double-slit experiment with diffraction of each slit included

In a double-slit experiment, the wavelength λ of the light source is 405 nm, the slit separation d is 19.44 μm , and the slit width a is 4.050 μm . Consider the interference of the light from the two slits and also the diffraction of the light through each slit.

- (a) How many bright interference fringes are within the central peak of the diffraction envelope?

KEY IDEAS

We first analyze the two basic mechanisms responsible for the optical pattern produced in the experiment:

- Single-slit diffraction:** The limits of the central peak are the first minima in the diffraction pattern due to either slit individually. (See Fig. 36-15.) The angular locations of those minima are given by Eq. 36-3 ($a \sin \theta = m_1 \lambda$). Here let us rewrite this equation as $a \sin \theta = m_1 \lambda$, with the subscript 1 referring to the one-slit diffraction. For the first minima in the diffraction pattern, we substitute $m_1 = 1$, obtaining

$$a \sin \theta = \lambda. \quad (36-22)$$

- Double-slit interference:** The angular locations of the bright fringes of the double-slit interference pattern are given by Eq. 35-14, which we can write as

$$d \sin \theta = m_2 \lambda, \quad \text{for } m_2 = 0, 1, 2, \dots \quad (36-23)$$

Here the subscript 2 refers to the double-slit interference.

Calculations: We can locate the first diffraction minimum within the double-slit fringe pattern by dividing Eq. 36-23 by Eq. 36-22 and solving for m_2 . By doing so and then substituting the given data, we obtain

$$m_2 = \frac{d}{a} = \frac{19.44 \mu\text{m}}{4.050 \mu\text{m}} = 4.8.$$

This tells us that the bright interference fringe for $m_2 = 4$ fits into the central peak of the one-slit diffraction pattern, but the fringe for $m_2 = 5$ does not fit. Within the central diffraction peak we have the central bright fringe ($m_2 = 0$), and four bright fringes (up to $m_2 = 4$) on each side of it. Thus, a total of nine bright fringes of the double-slit interference pattern are within the central peak of the diffraction envelope.

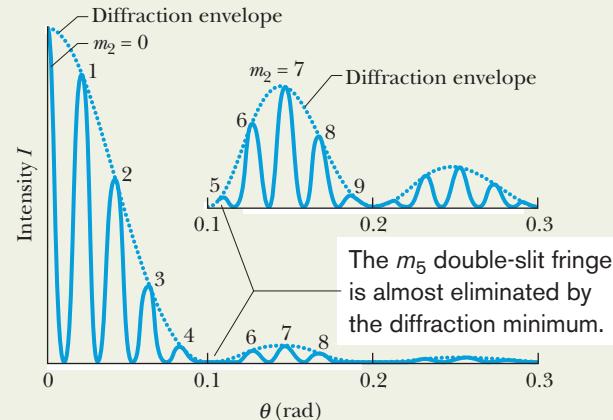


Fig. 36-17 One side of the intensity plot for a two-slit interference experiment. The inset shows (vertically expanded) the plot within the first and second side peaks of the diffraction envelope.

The bright fringes to one side of the central bright fringe are shown in Fig. 36-17.

- (b) How many bright fringes are within either of the first side peaks of the diffraction envelope?

KEY IDEA

The outer limits of the first side diffraction peaks are the second diffraction minima, each of which is at the angle θ given by $a \sin \theta = m_1 \lambda$ with $m_1 = 2$:

$$a \sin \theta = 2\lambda. \quad (36-24)$$

Calculation: Dividing Eq. 36-23 by Eq. 36-24, we find

$$m_2 = \frac{2d}{a} = \frac{(2)(19.44 \mu\text{m})}{4.050 \mu\text{m}} = 9.6.$$

This tells us that the second diffraction minimum occurs just before the bright interference fringe for $m_2 = 10$ in Eq. 36-23. Within either first side diffraction peak we have the fringes from $m_2 = 5$ to $m_2 = 9$, for a total of five bright fringes of the double-slit interference pattern (shown in the inset of Fig. 36-17). However, if the $m_2 = 5$ bright fringe, which is almost eliminated by the first diffraction minimum, is considered too dim to count, then only four bright fringes are in the first side diffraction peak.



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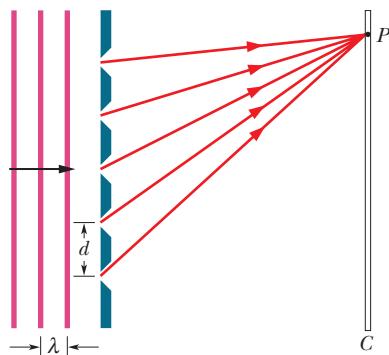


Fig. 36-18 An idealized diffraction grating, consisting of only five rulings, that produces an interference pattern on a distant viewing screen C .

36-8 Diffraction Gratings

One of the most useful tools in the study of light and of objects that emit and absorb light is the **diffraction grating**. This device is somewhat like the double-slit arrangement of Fig. 35-10 but has a much greater number N of slits, often called *rulings*, perhaps as many as several thousand per millimeter. An idealized grating consisting of only five slits is represented in Fig. 36-18. When monochromatic light is sent through the slits, it forms narrow interference fringes that can be analyzed to determine the wavelength of the light. (Diffraction gratings can also be opaque surfaces with narrow parallel grooves arranged like the slits in Fig. 36-18. Light then scatters back from the grooves to form interference fringes rather than being transmitted through open slits.)

With monochromatic light incident on a diffraction grating, if we gradually increase the number of slits from two to a large number N , the intensity plot changes from the typical double-slit plot of Fig. 36-15c to a much more complicated one and then eventually to a simple graph like that shown in Fig. 36-19a. The pattern you would see on a viewing screen using monochromatic red light from, say, a helium-neon laser is shown in Fig. 36-19b. The maxima are now very narrow (and so are called *lines*); they are separated by relatively wide dark regions.

We use a familiar procedure to find the locations of the bright lines on the viewing screen. We first assume that the screen is far enough from the grating so that the rays reaching a particular point P on the screen are approximately parallel when they leave the grating (Fig. 36-20). Then we apply to each pair of adjacent rulings the same reasoning we used for double-slit interference. The separation d between rulings is called the *grating spacing*. (If N rulings occupy a total width w , then $d = w/N$.) The path length difference between adjacent rays is again $d \sin \theta$ (Fig. 36-20), where θ is the angle from the central axis of the grating (and of the diffraction pattern) to point P . A line will be located at P if the path length difference between adjacent rays is an integer number of wavelengths—that is, if

$$d \sin \theta = m\lambda, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{maxima—lines}), \quad (36-25)$$

where λ is the wavelength of the light. Each integer m represents a different line; hence these integers can be used to label the lines, as in Fig. 36-19. The integers are then called the *order numbers*, and the lines are called the zeroth-order line (the central line, with $m = 0$), the first-order line ($m = 1$), the second-order line ($m = 2$), and so on.

If we rewrite Eq. 36-25 as $\theta = \sin^{-1}(m\lambda/d)$, we see that, for a given diffraction grating, the angle from the central axis to any line (say, the third-order line) depends on the wavelength of the light being used. Thus, when light of an unknown wavelength is sent through a diffraction grating, measurements of the angles to the higher-order lines can be used in Eq. 36-25 to determine the wavelength. Even light of several unknown wavelengths can be distinguished and identified in this way. We cannot do that with the double-slit arrangement of Section 35-4, even though the same equation and wavelength dependence apply there. In double-slit interference, the bright fringes due to different wavelengths overlap too much to be distinguished.

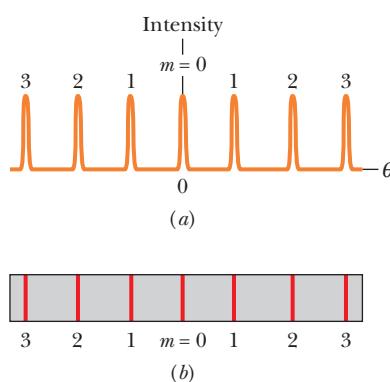


Fig. 36-19 (a) The intensity plot produced by a diffraction grating with a great many rulings consists of narrow peaks, here labeled with their order numbers m . (b) The corresponding bright fringes seen on the screen are called lines and are here also labeled with order numbers m .

Width of the Lines

A grating's ability to resolve (separate) lines of different wavelengths depends on the width of the lines. We shall here derive an expression for the *half-width* of the central line (the line for which $m = 0$) and then state an expression for the half-widths of the higher-order lines. We define the **half-width** of the central line as being the angle $\Delta\theta_{hw}$ from the center of the line at $\theta = 0$ outward to where the line effectively ends and darkness effectively begins with the first minimum

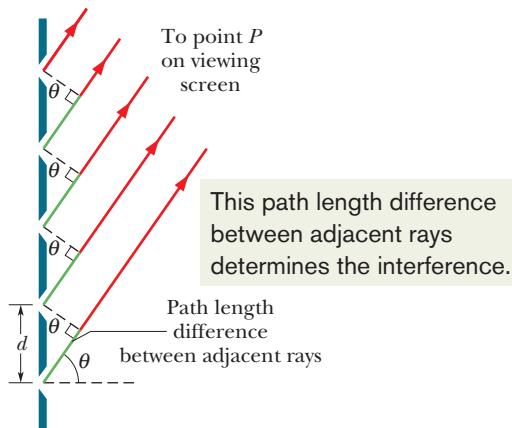


Fig. 36-20 The rays from the rulings in a diffraction grating to a distant point *P* are approximately parallel. The path length difference between each two adjacent rays is $d \sin \theta$, where θ is measured as shown. (The rulings extend into and out of the page.)

(Fig. 36-21). At such a minimum, the N rays from the N slits of the grating cancel one another. (The actual width of the central line is, of course, $2(\Delta\theta_{hw})$, but line widths are usually compared via half-widths.)

In Section 36-3 we were also concerned with the cancellation of a great many rays, there due to diffraction through a single slit. We obtained Eq. 36-3, which, because of the similarity of the two situations, we can use to find the first minimum here. It tells us that the first minimum occurs where the path length difference between the top and bottom rays equals λ . For single-slit diffraction, this difference is $a \sin \theta$. For a grating of N rulings, each separated from the next by distance d , the distance between the top and bottom rulings is Nd (Fig. 36-22), and so the path length difference between the top and bottom rays here is $Nd \sin \Delta\theta_{hw}$. Thus, the first minimum occurs where

$$Nd \sin \Delta\theta_{hw} = \lambda. \quad (36-26)$$

Because $\Delta\theta_{hw}$ is small, $\sin \Delta\theta_{hw} = \Delta\theta_{hw}$ (in radian measure). Substituting this in Eq. 36-26 gives the half-width of the central line as

$$\Delta\theta_{hw} = \frac{\lambda}{Nd} \quad (\text{half-width of central line}). \quad (36-27)$$

We state without proof that the half-width of any other line depends on its location relative to the central axis and is

$$\Delta\theta_{hw} = \frac{\lambda}{Nd \cos \theta} \quad (\text{half-width of line at } \theta). \quad (36-28)$$

Note that for light of a given wavelength λ and a given ruling separation d , the widths of the lines decrease with an increase in the number N of rulings. Thus, of two diffraction gratings, the grating with the larger value of N is better able to distinguish between wavelengths because its diffraction lines are narrower and so produce less overlap.

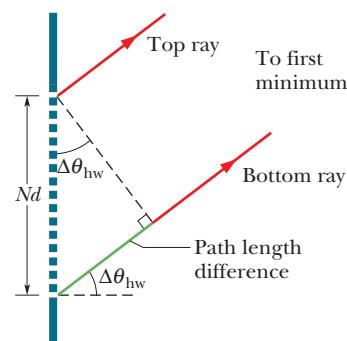


Fig. 36-22 The top and bottom rulings of a diffraction grating of N rulings are separated by Nd . The top and bottom rays passing through these rulings have a path length difference of $Nd \sin \Delta\theta_{hw}$, where $\Delta\theta_{hw}$ is the angle to the first minimum. (The angle is here greatly exaggerated for clarity.)

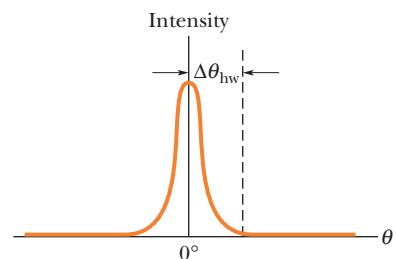


Fig. 36-21 The half-width $\Delta\theta_{hw}$ of the central line is measured from the center of that line to the adjacent minimum on a plot of I versus θ like Fig. 36-19a.

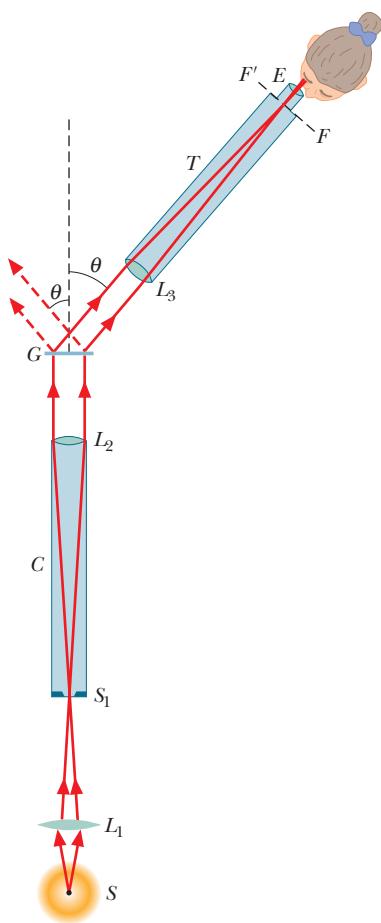


Fig. 36-23 A simple type of grating spectroscope used to analyze the wavelengths of light emitted by source S .

Grating Spectroscope

Diffraction gratings are widely used to determine the wavelengths that are emitted by sources of light ranging from lamps to stars. Figure 36-23 shows a simple *grating spectroscope* in which a grating is used for this purpose. Light from source S is focused by lens L_1 on a vertical slit S_1 placed in the focal plane of lens L_2 . The light emerging from tube C (called a *collimator*) is a plane wave and is incident perpendicularly on grating G , where it is diffracted into a diffraction pattern, with the $m = 0$ order diffracted at angle $\theta = 0$ along the central axis of the grating.

We can view the diffraction pattern that would appear on a viewing screen at any angle θ simply by orienting telescope T in Fig. 36-23 to that angle. Lens L_3 of the telescope then focuses the light diffracted at angle θ (and at slightly smaller and larger angles) onto a focal plane FF' within the telescope. When we look through eyepiece E , we see a magnified view of this focused image.

By changing the angle θ of the telescope, we can examine the entire diffraction pattern. For any order number other than $m = 0$, the original light is spread out according to wavelength (or color) so that we can determine, with Eq. 36-25, just what wavelengths are being emitted by the source. If the source emits discrete wavelengths, what we see as we rotate the telescope horizontally through the angles corresponding to an order m is a vertical line of color for each wavelength, with the shorter-wavelength line at a smaller angle θ than the longer-wavelength line.

For example, the light emitted by a hydrogen lamp, which contains hydrogen gas, has four discrete wavelengths in the visible range. If our eyes intercept this light directly, it appears to be white. If, instead, we view it through a grating spectroscope, we can distinguish, in several orders, the lines of the four colors corresponding to these visible wavelengths. (Such lines are called *emission lines*.) Four orders are represented in Fig. 36-24. In the central order ($m = 0$), the lines corresponding to all four wavelengths are superimposed, giving a single white line at $\theta = 0$. The colors are separated in the higher orders.

The third order is not shown in Fig. 36-24 for the sake of clarity; it actually overlaps the second and fourth orders. The fourth-order red line is missing because it is not formed by the grating used here. That is, when we attempt to solve Eq. 36-25 for the angle θ for the red wavelength when $m = 4$, we find that $\sin \theta$ is greater than unity, which is not possible. The fourth order is then said to be *incomplete* for this grating; it might not be incomplete for a grating with greater spacing d , which will spread the lines less than in Fig. 36-24. Figure 36-25 is a photograph of the visible emission lines produced by cadmium.

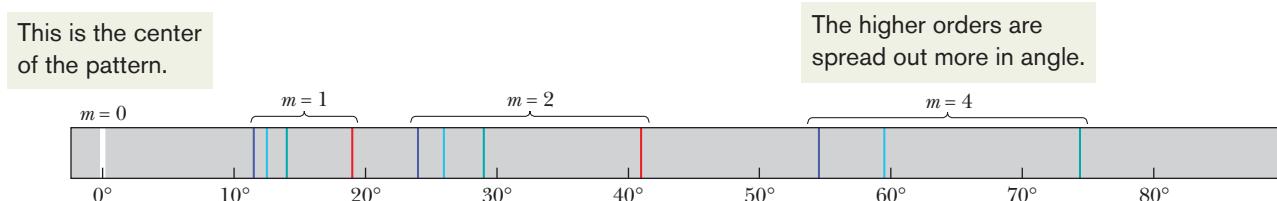


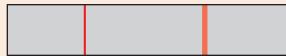
Fig. 36-24 The zeroth, first, second, and fourth orders of the visible emission lines from hydrogen. Note that the lines are farther apart at greater angles. (They are also dimmer and wider, although that is not shown here.)



Fig. 36-25 The visible emission lines of cadmium, as seen through a grating spectroscope. (Department of Physics, Imperial College/Science Photo Library/Photo Researchers)

**CHECKPOINT 5**

The figure shows lines of different orders produced by a diffraction grating in monochromatic red light. (a) Is the center of the pattern to the left or right? (b) In monochromatic green light, are the half-widths of the lines produced in the same orders greater than, less than, or the same as the half-widths of the lines shown?



36-9 Gratings: Dispersion and Resolving Power

Dispersion

To be useful in distinguishing wavelengths that are close to each other (as in a grating spectroscope), a grating must spread apart the diffraction lines associated with the various wavelengths. This spreading, called **dispersion**, is defined as

$$D = \frac{\Delta\theta}{\Delta\lambda} \quad (\text{dispersion defined}). \quad (36-29)$$

Here $\Delta\theta$ is the angular separation of two lines whose wavelengths differ by $\Delta\lambda$. The greater D is, the greater is the distance between two emission lines whose wavelengths differ by $\Delta\lambda$. We show below that the dispersion of a grating at angle θ is given by

$$D = \frac{m}{d \cos \theta} \quad (\text{dispersion of a grating}). \quad (36-30)$$

Thus, to achieve higher dispersion we must use a grating of smaller grating spacing d and work in a higher-order m . Note that the dispersion does not depend on the number of rulings N in the grating. The SI unit for D is the degree per meter or the radian per meter.



The fine rulings, each $0.5 \mu\text{m}$ wide, on a compact disc function as a diffraction grating. When a small source of white light illuminates a disc, the diffracted light forms colored “lunes” that are the composite of the diffraction patterns from the rulings.

(Kristen Brochmann/Fundamental Photographs)

Resolving Power

To *resolve* lines whose wavelengths are close together (that is, to make the lines distinguishable), the line should also be as narrow as possible. Expressed otherwise, the grating should have a high **resolving power** R , defined as

$$R = \frac{\lambda_{\text{avg}}}{\Delta\lambda} \quad (\text{resolving power defined}). \quad (36-31)$$

Here λ_{avg} is the mean wavelength of two emission lines that can barely be recognized as separate, and $\Delta\lambda$ is the wavelength difference between them. The greater R is, the closer two emission lines can be and still be resolved. We shall show below that the resolving power of a grating is given by the simple expression

$$R = Nm \quad (\text{resolving power of a grating}). \quad (36-32)$$

To achieve high resolving power, we must use many rulings (large N).

Proof of Eq. 36-30

Let us start with Eq. 36-25, the expression for the locations of the lines in the diffraction pattern of a grating:

$$d \sin \theta = m\lambda.$$

Let us regard θ and λ as variables and take differentials of this equation. We find

$$d(\cos \theta) d\theta = m d\lambda.$$

For small enough angles, we can write these differentials as small differences, obtaining

$$d(\cos \theta) \Delta\theta = m \Delta\lambda \quad (36-33)$$

or

$$\frac{\Delta\theta}{\Delta\lambda} = \frac{m}{d \cos \theta}.$$

The ratio on the left is simply D (see Eq. 36-29), and so we have indeed derived Eq. 36-30.

Proof of Eq. 36-32

We start with Eq. 36-33, which was derived from Eq. 36-25, the expression for the locations of the lines in the diffraction pattern formed by a grating. Here $\Delta\lambda$ is the small wavelength difference between two waves that are diffracted by the grating, and $\Delta\theta$ is the angular separation between them in the diffraction pattern. If $\Delta\theta$ is to be the smallest angle that will permit the two lines to be resolved, it must (by Rayleigh's criterion) be equal to the half-width of each line, which is given by Eq. 36-28:

$$\Delta\theta_{hw} = \frac{\lambda}{Nd \cos \theta}.$$

If we substitute $\Delta\theta_{hw}$ as given here for $\Delta\theta$ in Eq. 36-33, we find that

$$\frac{\lambda}{N} = m \Delta\lambda,$$

from which it readily follows that

$$R = \frac{\lambda}{\Delta\lambda} = Nm.$$

This is Eq. 36-32, which we set out to derive.

Dispersion and Resolving Power Compared

The resolving power of a grating must not be confused with its dispersion. Table 36-1 shows the characteristics of three gratings, all illuminated with light of wavelength $\lambda = 589$ nm, whose diffracted light is viewed in the first order ($m = 1$ in Eq. 36-25). You should verify that the values of D and R as given in the table can be calculated with Eqs. 36-30 and 36-32, respectively. (In the calculations for D , you will need to convert radians per meter to degrees per micrometer.)

For the conditions noted in Table 36-1, gratings A and B have the same dispersion D and A and C have the same resolving power R .

Figure 36-26 shows the intensity patterns (also called *line shapes*) that would be produced by these gratings for two lines of wavelengths λ_1 and λ_2 , in the vicinity of $\lambda = 589$ nm. Grating B , with the higher resolving power, produces narrower lines and thus is capable of distinguishing lines that are much closer together in wavelength than those in the figure. Grating C , with the higher dispersion, produces the greater angular separation between the lines.

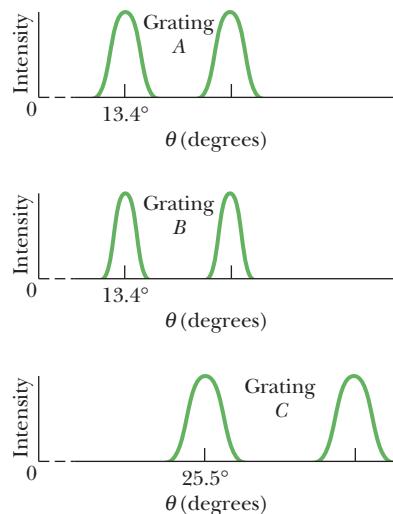


Fig. 36-26 The intensity patterns for light of two wavelengths sent through the gratings of Table 36-1. Grating B has the highest resolving power, and grating C the highest dispersion.

Table 36-1

Three Gratings^a

Grating	N	d (nm)	θ	D ($^{\circ}/\mu\text{m}$)	R
A	10 000	2540	13.4°	23.2	10 000
B	20 000	2540	13.4°	23.2	20 000
C	10 000	1360	25.5°	46.3	10 000

^aData are for $\lambda = 589$ nm and $m = 1$.

Sample Problem**Dispersion and resolving power of a diffraction grating**

A diffraction grating has 1.26×10^4 rulings uniformly spaced over width $w = 25.4$ mm. It is illuminated at normal incidence by yellow light from a sodium vapor lamp. This light contains two closely spaced emission lines (known as the sodium doublet) of wavelengths 589.00 nm and 589.59 nm.

- (a) At what angle does the first-order maximum occur (on either side of the center of the diffraction pattern) for the wavelength of 589.00 nm?

KEY IDEA

The maxima produced by the diffraction grating can be determined with Eq. 36-25 ($d \sin \theta = m\lambda$).

Calculations: The grating spacing d is

$$\begin{aligned} d &= \frac{w}{N} = \frac{25.4 \times 10^{-3} \text{ m}}{1.26 \times 10^4} \\ &= 2.016 \times 10^{-6} \text{ m} = 2016 \text{ nm}. \end{aligned}$$

The first-order maximum corresponds to $m = 1$. Substituting these values for d and m into Eq. 36-25 leads to

$$\begin{aligned} \theta &= \sin^{-1} \frac{m\lambda}{d} = \sin^{-1} \frac{(1)(589.00 \text{ nm})}{2016 \text{ nm}} \\ &= 16.99^\circ \approx 17.0^\circ. \quad (\text{Answer}) \end{aligned}$$

- (b) Using the dispersion of the grating, calculate the angular separation between the two lines in the first order.

KEY IDEAS

- (1) The angular separation $\Delta\theta$ between the two lines in the first order depends on their wavelength difference $\Delta\lambda$ and the dispersion D of the grating, according to Eq. 36-29 ($D = \Delta\theta/\Delta\lambda$). (2) The dispersion D depends on the angle θ at which it is to be evaluated.

Calculations: We can assume that, in the first order, the two sodium lines occur close enough to each other for us to



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evaluate D at the angle $\theta = 16.99^\circ$ we found in part (a) for one of those lines. Then Eq. 36-30 gives the dispersion as

$$\begin{aligned} D &= \frac{m}{d \cos \theta} = \frac{1}{(2016 \text{ nm})(\cos 16.99^\circ)} \\ &= 5.187 \times 10^{-4} \text{ rad/nm}. \end{aligned}$$

From Eq. 36-29 and with $\Delta\lambda$ in nanometers, we then have

$$\begin{aligned} \Delta\theta &= D \Delta\lambda = (5.187 \times 10^{-4} \text{ rad/nm})(589.59 - 589.00) \\ &= 3.06 \times 10^{-4} \text{ rad} = 0.0175^\circ. \quad (\text{Answer}) \end{aligned}$$

You can show that this result depends on the grating spacing d but not on the number of rulings there are in the grating.

- (c) What is the least number of rulings a grating can have and still be able to resolve the sodium doublet in the first order?

KEY IDEAS

- (1) The resolving power of a grating in any order m is physically set by the number of rulings N in the grating according to Eq. 36-32 ($R = Nm$). (2) The smallest wavelength difference $\Delta\lambda$ that can be resolved depends on the average wavelength involved and on the resolving power R of the grating, according to Eq. 36-31 ($R = \lambda_{\text{avg}}/\Delta\lambda$).

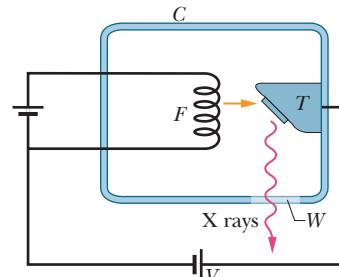
Calculation: For the sodium doublet to be barely resolved, $\Delta\lambda$ must be their wavelength separation of 0.59 nm, and λ_{avg} must be their average wavelength of 589.30 nm. Thus, we find that the smallest number of rulings for a grating to resolve the sodium doublet is

$$\begin{aligned} N &= \frac{R}{m} = \frac{\lambda_{\text{avg}}}{m \Delta\lambda} \\ &= \frac{589.30 \text{ nm}}{(1)(0.59 \text{ nm})} = 999 \text{ rulings}. \quad (\text{Answer}) \end{aligned}$$

36-10 X-Ray Diffraction

X rays are electromagnetic radiation whose wavelengths are of the order of 1 Å ($= 10^{-10} \text{ m}$). Compare this with a wavelength of 550 nm ($= 5.5 \times 10^{-7} \text{ m}$) at the center of the visible spectrum. Figure 36-27 shows that x rays are produced when electrons escaping from a heated filament F are accelerated by a potential difference V and strike a metal target T .

Fig. 36-27 X rays are generated when electrons leaving heated filament F are accelerated through a potential difference V and strike a metal target T . The “window” W in the evacuated chamber C is transparent to x rays.



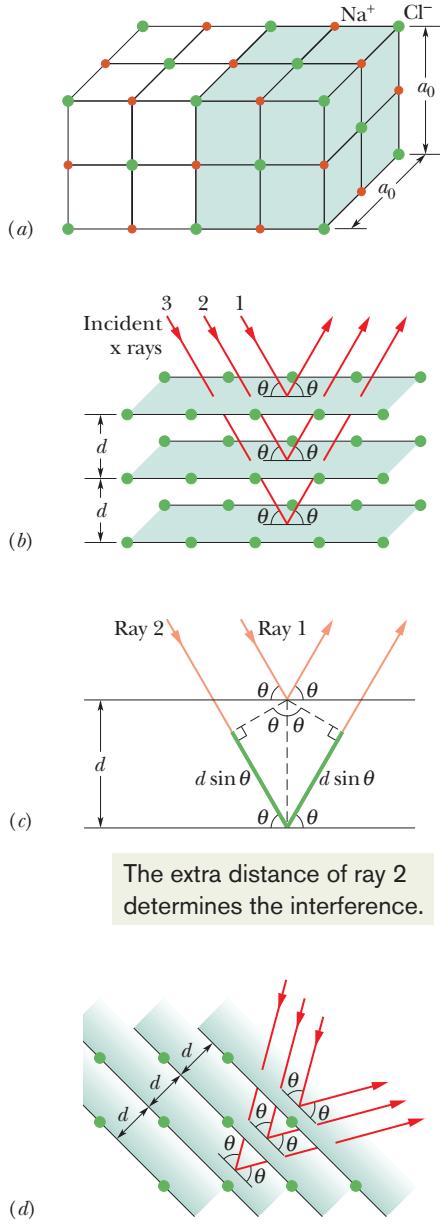


Fig. 36-28 (a) The cubic structure of NaCl, showing the sodium and chlorine ions and a unit cell (shaded). (b) Incident x rays undergo diffraction by the structure of (a). The x rays are diffracted as if they were reflected by a family of parallel planes, with the angle of reflection equal to the angle of incidence, both angles measured relative to the planes (not relative to a normal as in optics). (c) The path length difference between waves effectively reflected by two adjacent planes is $2d \sin \theta$. (d) A different orientation of the incident x rays relative to the structure. A different family of parallel planes now effectively reflects the x rays.

A standard optical diffraction grating cannot be used to discriminate between different wavelengths in the x-ray wavelength range. For $\lambda = 1 \text{ \AA}$ ($= 0.1 \text{ nm}$) and $d = 3000 \text{ nm}$, for example, Eq. 36-25 shows that the first-order maximum occurs at

$$\theta = \sin^{-1} \frac{m\lambda}{d} = \sin^{-1} \frac{(1)(0.1 \text{ nm})}{3000 \text{ nm}} = 0.0019^\circ.$$

This is too close to the central maximum to be practical. A grating with $d \approx \lambda$ is desirable, but, because x-ray wavelengths are about equal to atomic diameters, such gratings cannot be constructed mechanically.

In 1912, it occurred to German physicist Max von Laue that a crystalline solid, which consists of a regular array of atoms, might form a natural three-dimensional “diffraction grating” for x rays. The idea is that, in a crystal such as sodium chloride (NaCl), a basic unit of atoms (called the *unit cell*) repeats itself throughout the array. Figure 36-28a represents a section through a crystal of NaCl and identifies this basic unit. The unit cell is a cube measuring a_0 on each side.

When an x-ray beam enters a crystal such as NaCl, x rays are *scattered*—that is, redirected—in all directions by the crystal structure. In some directions the scattered waves undergo destructive interference, resulting in intensity minima; in other directions the interference is constructive, resulting in intensity maxima. This process of scattering and interference is a form of diffraction.

Although the process of diffraction of x rays by a crystal is complicated, the maxima turn out to be in directions *as if* the x rays were reflected by a family of parallel *reflecting planes* (or *crystal planes*) that extend through the atoms within the crystal and that contain regular arrays of the atoms. (The x rays are not actually reflected; we use these fictional planes only to simplify the analysis of the actual diffraction process.)

Figure 36-28b shows three reflecting planes (part of a family containing many parallel planes) with *interplanar spacing* d , from which the incident rays shown are said to reflect. Rays 1, 2, and 3 reflect from the first, second, and third planes, respectively. At each reflection the angle of incidence and the angle of reflection are represented with θ . Contrary to the custom in optics, these angles are defined relative to the *surface* of the reflecting plane rather than a normal to that surface. For the situation of Fig. 36-28b, the interplanar spacing happens to be equal to the unit cell dimension a_0 .

Figure 36-28c shows an edge-on view of reflection from an adjacent pair of planes. The waves of rays 1 and 2 arrive at the crystal in phase. After they are reflected, they must again be in phase because the reflections and the reflecting planes have been defined solely to explain the intensity maxima in the diffraction of x rays by a crystal. Unlike light rays, the x rays do not refract upon entering the crystal; moreover, we do not define an index of refraction for this situation. Thus, the relative phase between the waves of rays 1 and 2 as they leave the crystal is set solely by their path length difference. For these rays to be in phase, the path length difference must be equal to an integer multiple of the wavelength λ of the x rays.

By drawing the dashed perpendiculars in Fig. 36-28c, we find that the path length difference is $2d \sin \theta$. In fact, this is true for any pair of adjacent planes in the family of planes represented in Fig. 36-28b. Thus, we have, as the criterion for intensity maxima for x-ray diffraction,

$$2d \sin \theta = m\lambda, \quad \text{for } m = 1, 2, 3, \dots \quad (\text{Bragg's law}), \quad (36-34)$$

where m is the order number of an intensity maximum. Equation 36-34 is called **Bragg's law** after British physicist W. L. Bragg, who first derived it. (He and his father shared the 1915 Nobel Prize in physics for their use of x rays to study the structures of crystals.) The angle of incidence and reflection in Eq. 36-34 is called a *Bragg angle*.

Regardless of the angle at which x rays enter a crystal, there is always a family of planes from which they can be said to reflect so that we can apply Bragg's law. In Fig. 36-28d, notice that the crystal structure has the same orientation as it does in Fig. 36-28a, but the angle at which the beam enters the structure differs

from that shown in Fig. 36-28b. This new angle requires a new family of reflecting planes, with a different interplanar spacing d and different Bragg angle θ , in order to explain the x-ray diffraction via Bragg's law.

Figure 36-29 shows how the interplanar spacing d can be related to the unit cell dimension a_0 . For the particular family of planes shown there, the Pythagorean theorem gives

$$5d = \sqrt{\frac{5}{4}a_0^2},$$

or

$$d = \frac{a_0}{\sqrt{20}} = 0.2236a_0. \quad (36-35)$$

Figure 36-29 suggests how the dimensions of the unit cell can be found once the interplanar spacing has been measured by means of x-ray diffraction.

X-ray diffraction is a powerful tool for studying both x-ray spectra and the arrangement of atoms in crystals. To study spectra, a particular set of crystal planes, having a known spacing d , is chosen. These planes effectively reflect different wavelengths at different angles. A detector that can discriminate one angle from another can then be used to determine the wavelength of radiation reaching it. The crystal itself can be studied with a monochromatic x-ray beam, to determine not only the spacing of various crystal planes but also the structure of the unit cell.

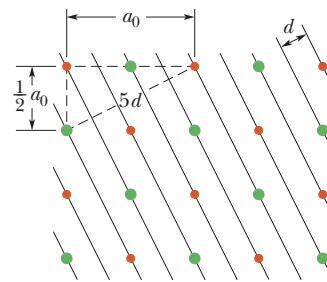


Fig. 36-29 A family of planes through the structure of Fig. 36-28a, and a way to relate the edge length a_0 of a unit cell to the interplanar spacing d .

REVIEW & SUMMARY

Diffraction When waves encounter an edge, an obstacle, or an aperture the size of which is comparable to the wavelength of the waves, those waves spread out as they travel and, as a result, undergo interference. This is called **diffraction**.

Single-Slit Diffraction Waves passing through a long narrow slit of width a produce, on a viewing screen, a **single-slit diffraction pattern** that includes a central maximum and other maxima, separated by minima located at angles θ to the central axis that satisfy

$$a \sin \theta = m\lambda, \quad \text{for } m = 1, 2, 3, \dots \quad (\text{minima}). \quad (36-3)$$

The intensity of the diffraction pattern at any given angle θ is

$$I(\theta) = I_m \left(\frac{\sin \alpha}{\alpha} \right)^2, \quad \text{where } \alpha = \frac{\pi a}{\lambda} \sin \theta \quad (36-5, 36-6)$$

and I_m is the intensity at the center of the pattern.

Circular-Aperture Diffraction Diffraction by a circular aperture or a lens with diameter d produces a central maximum and concentric maxima and minima, with the first minimum at an angle θ given by

$$\sin \theta = 1.22 \frac{\lambda}{d} \quad (\text{first minimum—circular aperture}). \quad (36-12)$$

Rayleigh's Criterion Rayleigh's criterion suggests that two objects are on the verge of resolvability if the central diffraction maximum of one is at the first minimum of the other. Their angular separation must then be at least

$$\theta_R = 1.22 \frac{\lambda}{d} \quad (\text{Rayleigh's criterion}), \quad (36-14)$$

in which d is the diameter of the aperture through which the light passes.

Double-Slit Diffraction Waves passing through two slits, each of width a , whose centers are a distance d apart, display diffraction patterns whose intensity I at angle θ is

$$I(\theta) = I_m (\cos^2 \beta) \left(\frac{\sin \alpha}{\alpha} \right)^2 \quad (\text{double slit}), \quad (36-19)$$

with $\beta = (\pi d / \lambda) \sin \theta$ and α as for single-slit diffraction.

Diffraction Gratings A *diffraction grating* is a series of "slits" used to separate an incident wave into its component wavelengths by separating and displaying their diffraction maxima. Diffraction by N (multiple) slits results in maxima (lines) at angles θ such that

$$d \sin \theta = m\lambda, \quad \text{for } m = 0, 1, 2, \dots \quad (\text{maxima}), \quad (36-25)$$

with the **half-widths** of the lines given by

$$\Delta\theta_{hw} = \frac{\lambda}{Nd \cos \theta} \quad (\text{half-widths}). \quad (36-28)$$

The dispersion D and resolving power R are given by

$$D = \frac{\Delta\theta}{\Delta\lambda} = \frac{m}{d \cos \theta} \quad (36-29, 36-30)$$

and

$$R = \frac{\lambda_{avg}}{\Delta\lambda} = Nm. \quad (36-31, 36-32)$$

X-Ray Diffraction The regular array of atoms in a crystal is a three-dimensional diffraction grating for short-wavelength waves such as x rays. For analysis purposes, the atoms can be visualized as being arranged in planes with characteristic interplanar spacing d . Diffraction maxima (due to constructive interference) occur if the incident direction of the wave, measured from the surfaces of these planes, and the wavelength λ of the radiation satisfy **Bragg's law**:

$$2d \sin \theta = m\lambda, \quad \text{for } m = 1, 2, 3, \dots \quad (\text{Bragg's law}). \quad (36-34)$$

Q U E S T I O N S

1 You are conducting a single-slit diffraction experiment with light of wavelength λ . What appears, on a distant viewing screen, at a point at which the top and bottom rays through the slit have a path length difference equal to (a) 5λ and (b) 4.5λ ?

2 In a single-slit diffraction experiment, the top and bottom rays through the slit arrive at a certain point on the viewing screen with a path length difference of 4.0 wavelengths. In a phasor representation like those in Fig 36-7, how many overlapping circles does the chain of phasors make?

3 For three experiments, Fig. 36-30 gives the parameter β of Eq. 36-20 versus angle θ for two-slit interference using light of wavelength 500 nm. The slit separations in the three experiments differ. Rank the experiments according to (a) the slit separations and (b) the total number of two-slit interference maxima in the pattern, greatest first.

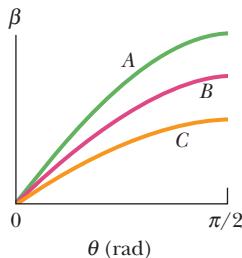


Fig. 36-30
Question 3.

4 For three experiments, Fig. 36-31 gives α versus angle θ in one-slit diffraction using light of wavelength 500 nm. Rank the experiments according to (a) the slit widths and (b) the total number of diffraction minima in the pattern, greatest first.

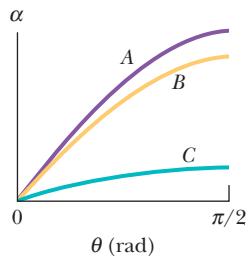


Fig. 36-31 Question 4.

5 Figure 36-32 shows four choices for the rectangular opening of a source of either sound waves or light waves. The sides have lengths of either L or $2L$, with L being 3.0 times the wavelength of the waves. Rank the openings according to the extent of (a) left-right spreading and (b) up-down spreading of the waves due to diffraction, greatest first.

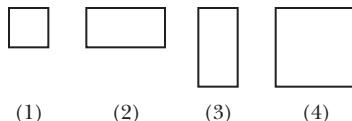


Fig. 36-32 Question 5.

6 Light of frequency f illuminating a long narrow slit produces a diffraction pattern. (a) If we switch to light of frequency $1.3f$, does the pattern expand away from the center or contract toward the center? (b) Does the pattern expand or contract if, instead, we submerge the equipment in clear corn syrup?

7 At night many people see rings (called *entoptic halos*) surrounding bright outdoor lamps in otherwise dark surroundings. The rings are the first of the side maxima in diffraction patterns

produced by structures that are thought to be within the cornea (or possibly the lens) of the observer's eye. (The central maxima of such patterns overlap the lamp.) (a) Would a particular ring become smaller or larger if the lamp were switched from blue to red light? (b) If a lamp emits white light, is blue or red on the outside edge of the ring?

8 (a) For a given diffraction grating, does the smallest difference $\Delta\lambda$ in two wavelengths that can be resolved increase, decrease, or remain the same as the wavelength increases? (b) For a given wavelength region (say, around 500 nm), is $\Delta\lambda$ greater in the first order or in the third order?

9 Figure 36-33 shows a red line and a green line of the same order in the pattern produced by a diffraction grating. If we increased the number of rulings in the grating—say, by removing tape that had covered the outer half of the rulings—would (a) the half-widths of the lines and (b) the separation of the lines increase, decrease, or remain the same? (c) Would the lines shift to the right, shift to the left, or remain in place?



Fig. 36-33 Questions 9 and 10.

10 For the situation of Question 9 and Fig. 36-33, if instead we increased the grating spacing, would (a) the half-widths of the lines and (b) the separation of the lines increase, decrease, or remain the same? (c) Would the lines shift to the right, shift to the left, or remain in place?

11 (a) Figure 36-34a shows the lines produced by diffraction gratings *A* and *B* using light of the same wavelength; the lines are of the same order and appear at the same angles θ . Which grating has the greater number of rulings? (b) Figure 36-34b shows lines of two orders produced by a single diffraction grating using light of two wavelengths, both in the red region of the spectrum. Which lines, the left pair or right pair, are in the order with greater m ? Is the center of the diffraction pattern located to the left or to the right in (c) Fig. 36-34a and (d) Fig. 36-34b?

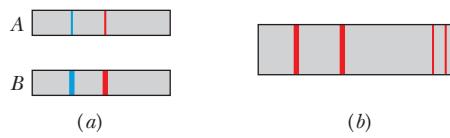


Fig. 36-34 Question 11.

12 Figure 36-35 shows the bright fringes that lie within the central diffraction envelope in two double-slit diffraction experiments using the same wavelength of light. Are (a) the slit width a , (b) the slit separation d , and (c) the ratio d/a in experiment *B* greater than, less than, or the same as those quantities in experiment *A*?



Fig. 36-35 Question 12.



PROBLEMS



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty

Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>**sec. 36-3 Diffraction by a Single Slit: Locating the Minima**

- 1 GO** The distance between the first and fifth minima of a single-slit diffraction pattern is 0.35 mm with the screen 40 cm away from the slit, when light of wavelength 550 nm is used. (a) Find the slit width. (b) Calculate the angle θ of the first diffraction minimum.

- 2** What must be the ratio of the slit width to the wavelength for a single slit to have the first diffraction minimum at $\theta = 45.0^\circ$?

- 3** A plane wave of wavelength 590 nm is incident on a slit with a width of $a = 0.40$ mm. A thin converging lens of focal length +70 cm is placed between the slit and a viewing screen and focuses the light on the screen. (a) How far is the screen from the lens? (b) What is the distance on the screen from the center of the diffraction pattern to the first minimum?

- 4** In conventional television, signals are broadcast from towers to home receivers. Even when a receiver is not in direct view of a tower because of a hill or building, it can still intercept a signal if the signal diffracts enough around the obstacle, into the obstacle's "shadow region." Previously, television signals had a wavelength of about 50 cm, but digital television signals that are transmitted from towers have a wavelength of about 10 mm. (a) Did this change in wavelength increase or decrease the diffraction of the signals into the shadow regions of obstacles? Assume that a signal passes through an opening of 5.0 m width between two adjacent buildings. What is the angular spread of the central diffraction maximum (out to the first minima) for wavelengths of (b) 50 cm and (c) 10 mm?

- 5** A single slit is illuminated by light of wavelengths λ_a and λ_b , chosen so that the first diffraction minimum of the λ_a component coincides with the second minimum of the λ_b component. (a) If $\lambda_b = 350$ nm, what is λ_a ? For what order number m_b (if any) does a minimum of the λ_b component coincide with the minimum of the λ_a component in the order number (b) $m_a = 2$ and (c) $m_a = 3$?

- 6** Monochromatic light of wavelength 441 nm is incident on a narrow slit. On a screen 2.00 m away, the distance between the second diffraction minimum and the central maximum is 1.50 cm. (a) Calculate the angle of diffraction θ of the second minimum. (b) Find the width of the slit.

- 7** Light of wavelength 633 nm is incident on a narrow slit. The angle between the first diffraction minimum on one side of the central maximum and the first minimum on the other side is 1.20° . What is the width of the slit?

- 8** Sound waves with frequency 3000 Hz and speed 343 m/s diffract through the rectangular opening of a speaker cabinet and into a large auditorium of length $d = 100$ m. The opening, which has a horizontal width of 30.0 cm, faces a wall 100 m away (Fig. 36-36). Along that wall, how far from the central axis will a

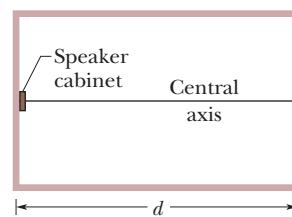


Fig. 36-36 Problem 8.

listener be at the first diffraction minimum and thus have difficulty hearing the sound? (Neglect reflections.)

- 9 SSM ILW** A slit 1.00 mm wide is illuminated by light of wavelength 589 nm. We see a diffraction pattern on a screen 3.00 m away. What is the distance between the first two diffraction minima on the same side of the central diffraction maximum?

- 10** Manufacturers of wire (and other objects of small dimension) sometimes use a laser to continually monitor the thickness of the product. The wire intercepts the laser beam, producing a diffraction pattern like that of a single slit of the same width as the wire diameter (Fig. 36-37). Suppose a helium-neon laser, of wavelength 632.8 nm, illuminates a wire, and the diffraction pattern appears on a screen at distance $L = 2.60$ m. If the desired wire diameter is 1.37 mm, what is the observed distance between the two tenth-order minima (one on each side of the central maximum)?

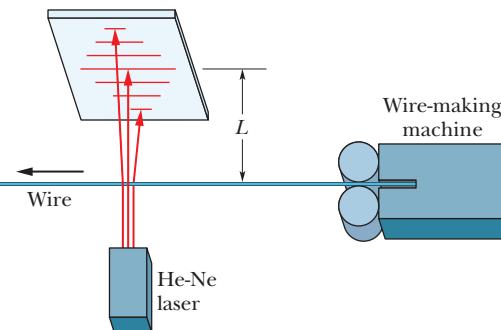


Fig. 36-37 Problem 10.

sec. 36-5 Intensity in Single-Slit Diffraction, Quantitatively

- 11** A 0.10-mm-wide slit is illuminated by light of wavelength 589 nm. Consider a point P on a viewing screen on which the diffraction pattern of the slit is viewed; the point is at 30° from the central axis of the slit. What is the phase difference between the Huygens wavelets arriving at point P from the top and midpoint of the slit? (Hint: See Eq. 36-4.)

- 12** Figure 36-38 gives α versus the sine of the angle θ in a single-slit diffraction experiment using light of wavelength 610 nm. The vertical axis scale is set by $\alpha_s = 12$ rad. What are (a) the slit width, (b) the total number of diffraction minima in the pattern (count them on both sides of the center of the diffraction pattern), (c) the least angle for a minimum, and (d) the greatest angle for a minimum?

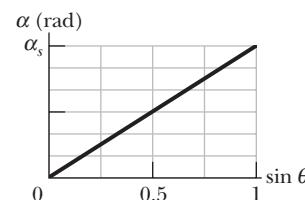


Fig. 36-38 Problem 12.

- 13** Monochromatic light with wavelength 538 nm is incident on a slit with width 0.025 mm. The distance from the slit to a screen is 3.5 m. Consider a point on the screen 1.1 cm from the central maximum. Calculate (a) θ for that point, (b) α , and (c) the ratio of the intensity at that point to the intensity at the central maximum.

1016 CHAPTER 36 DIFFRACTION

•14 In the single-slit diffraction experiment of Fig. 36-4, let the wavelength of the light be 500 nm, the slit width be 6.00 μm , and the viewing screen be at distance $D = 3.00 \text{ m}$. Let a y axis extend upward along the viewing screen, with its origin at the center of the diffraction pattern. Also let I_P represent the intensity of the diffracted light at point P at $y = 15.0 \text{ cm}$. (a) What is the ratio of I_P to the intensity I_m at the center of the pattern? (b) Determine where point P is in the diffraction pattern by giving the maximum and minimum between which it lies, or the two minima between which it lies.

•15 **SSM WWW** The full width at half-maximum (FWHM) of a central diffraction maximum is defined as the angle between the two points in the pattern where the intensity is one-half that at the center of the pattern. (See Fig. 36-8b.) (a) Show that the intensity drops to one-half the maximum value when $\sin^2 \alpha = \alpha^2/2$. (b) Verify that $\alpha = 1.39 \text{ rad}$ (about 80°) is a solution to the transcendental equation of (a). (c) Show that the FWHM is $\Delta\theta = 2 \sin^{-1}(0.443\lambda/a)$, where a is the slit width. Calculate the FWHM of the central maximum for slit width (d) 1.00λ , (e) 5.00λ , and (f) 10.0λ .

•16 *Babinet's principle.* A monochromatic beam of parallel light is incident on a “collimating” hole of diameter $x \gg \lambda$. Point P lies in the geometrical shadow region on a *distant* screen (Fig. 36-39a). Two diffracting objects, shown in Fig. 36-39b, are placed in turn over the collimating hole. Object A is an opaque circle with a hole in it, and B is the “photographic negative” of A . Using superposition concepts, show that the intensity at P is identical for the two diffracting objects A and B .

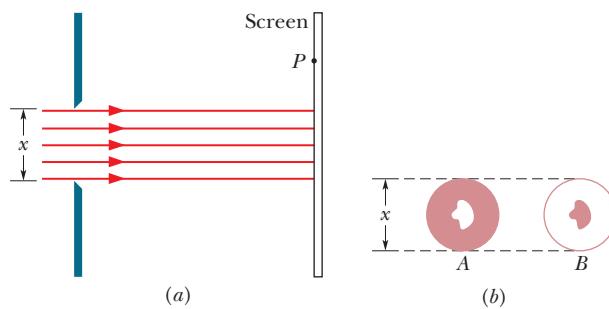


Fig. 36-39 Problem 16.

•17 (a) Show that the values of α at which intensity maxima for single-slit diffraction occur can be found exactly by differentiating Eq. 36-5 with respect to α and equating the result to zero, obtaining the condition $\tan \alpha = \alpha$. To find values of α satisfying this relation, plot the curve $y = \tan \alpha$ and the straight line $y = \alpha$ and then find their intersections, or use a calculator to find an appropriate value of α by trial and error. Next, from $\alpha = (m + \frac{1}{2})\pi$, determine the values of m associated with the maxima in the single-slit pattern. (These m values are *not* integers because secondary maxima do not lie exactly halfway between minima.) What are the (b) smallest α and (c) associated m , the (d) second smallest α and (e) associated m , and the (f) third smallest α and (g) associated m ?

sec. 36-6 Diffraction by a Circular Aperture

•18 The wall of a large room is covered with acoustic tile in which small holes are drilled 5.0 mm from center to center. How far can a person be from such a tile and still distinguish the individual holes, assuming ideal conditions, the pupil diameter of the observer's eye to be 4.0 mm, and the wavelength of the room light to be 550 nm?

•19 (a) How far from grains of red sand must you be to position yourself just at the limit of resolving the grains if your pupil diameter is 1.5 mm, the grains are spherical with radius 50 μm , and the light from the grains has wavelength 650 nm? (b) If the grains were blue and the light from them had wavelength 400 nm, would the answer to (a) be larger or smaller?

•20 The radar system of a navy cruiser transmits at a wavelength of 1.6 cm, from a circular antenna with a diameter of 2.3 m. At a range of 6.2 km, what is the smallest distance that two speedboats can be from each other and still be resolved as two separate objects by the radar system?

•21 **SSM WWW** Estimate the linear separation of two objects on Mars that can just be resolved under ideal conditions by an observer on Earth (a) using the naked eye and (b) using the 200 in. (= 5.1 m) Mount Palomar telescope. Use the following data: distance to Mars = $8.0 \times 10^7 \text{ km}$, diameter of pupil = 5.0 mm, wavelength of light = 550 nm.

•22 Assume that Rayleigh's criterion gives the limit of resolution of an astronaut's eye looking down on Earth's surface from a typical space shuttle altitude of 400 km. (a) Under that idealized assumption, estimate the smallest linear width on Earth's surface that the astronaut can resolve. Take the astronaut's pupil diameter to be 5 mm and the wavelength of visible light to be 550 nm. (b) Can the astronaut resolve the Great Wall of China (Fig. 36-40), which is more than 3000 km long, 5 to 10 m thick at its base, 4 m thick at its top, and 8 m in height? (c) Would the astronaut be able to resolve any unmistakable sign of intelligent life on Earth's surface?



Fig. 36-40 Problem 22. The Great Wall of China.
(AP/Wide World Photos)

•23 **SSM** The two headlights of an approaching automobile are 1.4 m apart. At what (a) angular separation and (b) maximum distance will the eye resolve them? Assume that the pupil diameter is 5.0 mm, and use a wavelength of 550 nm for the light. Also assume that diffraction effects alone limit the resolution so that Rayleigh's criterion can be applied.

•24 *Entoptic halos.* If someone looks at a bright outdoor lamp in otherwise dark surroundings, the lamp appears to be surrounded by bright and dark rings (hence *halos*) that are actually a circular diffraction pattern as in Fig. 36-10, with the central maximum overlapping the direct light from the lamp. The diffraction is produced by structures within the cornea or lens of the eye (hence *entop-*

tic). If the lamp is monochromatic at wavelength 550 nm and the first dark ring subtends angular diameter 2.5° in the observer's view, what is the (linear) diameter of the structure producing the diffraction?

•25 ILW Find the separation of two points on the Moon's surface that can just be resolved by the 200 in. (= 5.1 m) telescope at Mount Palomar, assuming that this separation is determined by diffraction effects. The distance from Earth to the Moon is 3.8×10^5 km. Assume a wavelength of 550 nm for the light.

•26 The telescopes on some commercial surveillance satellites can resolve objects on the ground as small as 85 cm across (see Google Earth), and the telescopes on military surveillance satellites reportedly can resolve objects as small as 10 cm across. Assume first that object resolution is determined entirely by Rayleigh's criterion and is not degraded by turbulence in the atmosphere. Also assume that the satellites are at a typical altitude of 400 km and that the wavelength of visible light is 550 nm. What would be the required diameter of the telescope aperture for (a) 85 cm resolution and (b) 10 cm resolution? (c) Now, considering that turbulence is certain to degrade resolution and that the aperture diameter of the Hubble Space Telescope is 2.4 m, what can you say about the answer to (b) and about how the military surveillance resolutions are accomplished?

•27 If Superman really had x-ray vision at 0.10 nm wavelength and a 4.0 mm pupil diameter, at what maximum altitude could he distinguish villains from heroes, assuming that he needs to resolve points separated by 5.0 cm to do this?

•28 GO The wings of tiger beetles (Fig. 36-41) are colored by interference due to thin cuticle-like layers. In addition, these layers are arranged in patches that are $60\ \mu\text{m}$ across and produce different colors. The color you see is a pointillistic mixture of thin-film interference colors that varies with perspective. Approximately what viewing distance from a wing puts you at the limit of resolving the different colored patches according to Rayleigh's criterion? Use 550 nm as the wavelength of light and 3.00 mm as the diameter of your pupil.



Fig. 36-41 Problem 28. Tiger beetles are colored by pointillistic mixtures of thin-film interference colors.
(Kjell B. Sandved/Bruce Coleman, Inc.)

•29 (a) What is the angular separation of two stars if their images are barely resolved by the Thaw refracting telescope at the Allegheny Observatory in Pittsburgh? The lens diameter is 76 cm and its focal length is 14 m. Assume $\lambda = 550$ nm. (b) Find the distance between these barely resolved stars if each of them is 10 light-years distant from Earth. (c) For the image of a single star in this telescope, find the diameter of the first dark ring in the diffraction pattern, as measured on a photographic plate placed at the focal plane of the telescope lens. Assume that the structure of the image is associated entirely with diffraction at the lens aperture and not with lens "errors."

•30 GO **Floating dots.** The floaters you see when viewing a bright, featureless background are diffraction patterns of defects in the vitreous humor that fills most of your eye. Sighting through a pinhole sharpens the diffraction pattern. If you also view a small circular dot, you can approximate the defect's size. Assume that the defect diffracts light as a circular aperture does. Adjust the dot's distance L from your eye (or eye lens) until the dot and the circle of the first minimum in the diffraction pattern appear to have the same size in your view. That is, until they have the same diameter D' on the retina at distance $L' = 2.0$ cm from the front of the eye, as suggested in Fig. 36-42a, where the angles on the two sides of the eye lens are equal. Assume that the wavelength of visible light is $\lambda = 550$ nm. If the dot has diameter $D = 2.0$ mm and is distance $L = 45.0$ cm from the eye and the defect is $x = 6.0$ mm in front of the retina (Fig. 36-42b), what is the diameter of the defect?

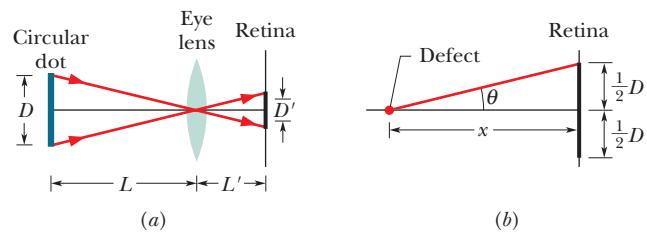


Fig. 36-42 Problem 30.

•31 SSM Millimeter-wave radar generates a narrower beam than conventional microwave radar, making it less vulnerable to antiradar missiles than conventional radar. (a) Calculate the angular width 2θ of the central maximum, from first minimum to first minimum, produced by a 220 GHz radar beam emitted by a 55.0-cm-diameter circular antenna. (The frequency is chosen to coincide with a low-absorption atmospheric "window.") (b) What is 2θ for a more conventional circular antenna that has a diameter of 2.3 m and emits at wavelength 1.6 cm?

•32 (a) A circular diaphragm 60 cm in diameter oscillates at a frequency of 25 kHz as an underwater source of sound used for submarine detection. Far from the source, the sound intensity is distributed as the diffraction pattern of a circular hole whose diameter equals that of the diaphragm. Take the speed of sound in water to be 1450 m/s and find the angle between the normal to the diaphragm and a line from the diaphragm to the first minimum. (b) Is there such a minimum for a source having an (audible) frequency of 1.0 kHz?

•33 GO Nuclear-pumped x-ray lasers are seen as a possible weapon to destroy ICBM booster rockets at ranges up to 2000 km. One limitation on such a device is the spreading of the beam due to diffraction, with resulting dilution of beam intensity. Consider such

1018 CHAPTER 36 DIFFRACTION

a laser operating at a wavelength of 1.40 nm. The element that emits light is the end of a wire with diameter 0.200 mm. (a) Calculate the diameter of the central beam at a target 2000 km away from the beam source. (b) By what factor is the beam intensity reduced in transit to the target? (The laser is fired from space, so that atmospheric absorption can be ignored.)

••34  A circular obstacle produces the same diffraction pattern as a circular hole of the same diameter (except very near $\theta = 0$). Airborne water drops are examples of such obstacles. When you see the Moon through suspended water drops, such as in a fog, you intercept the diffraction pattern from many drops. The composite of the central diffraction maxima of those drops forms a white region that surrounds the Moon and may obscure it. Figure 36-43 is a photograph in which the Moon is obscured. There are two faint, colored rings around the Moon (the larger one may be too faint to be seen in your copy of the photograph). The smaller ring is on the outer edge of the central maxima from the drops; the somewhat larger ring is on the outer edge of the smallest of the secondary maxima from the drops (see Fig. 36-10). The color is visible because the rings are adjacent to the diffraction minima (dark rings) in the patterns. (Colors in other parts of the pattern overlap too much to be visible.)

(a) What is the color of these rings on the outer edges of the diffraction maxima? (b) The colored ring around the central maxima in Fig. 36-43 has an angular diameter that is 1.35 times the angular diameter of the Moon, which is 0.50° . Assume that the drops all have about the same diameter. Approximately what is that diameter?



Fig. 36-43 Problem 34. The corona around the Moon is a composite of the diffraction patterns of airborne water drops. (Pekka Parvinen/Photo Researchers)

sec. 36-7 Diffraction by a Double Slit

•35 Suppose that the central diffraction envelope of a double-slit diffraction pattern contains 11 bright fringes and the first diffraction minima eliminate (are coincident with) bright fringes. How many bright fringes lie between the first and second minima of the diffraction envelope?

•36 A beam of light of a single wavelength is incident perpendicularly on a double-slit arrangement, as in Fig. 35-10. The slit widths

are each $46 \mu\text{m}$ and the slit separation is 0.30 mm. How many complete bright fringes appear between the two first-order minima of the diffraction pattern?

•37 In a double-slit experiment, the slit separation d is 2.00 times the slit width w . How many bright interference fringes are in the central diffraction envelope?

•38 In a certain two-slit interference pattern, 10 bright fringes lie within the second side peak of the diffraction envelope and diffraction minima coincide with two-slit interference maxima. What is the ratio of the slit separation to the slit width?

•39 Light of wavelength 440 nm passes through a double slit, yielding a diffraction pattern whose graph of intensity I versus angular position θ is shown in Fig. 36-44. Calculate (a) the slit width and (b) the slit separation. (c) Verify the displayed intensities of the $m = 1$ and $m = 2$ interference fringes.

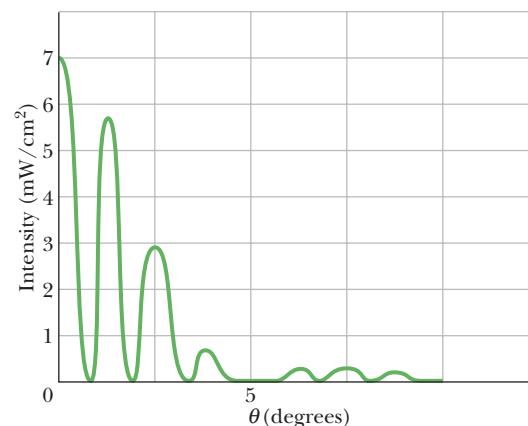


Fig. 36-44 Problem 39.

•40 Figure 36-45 gives the parameter β of Eq. 36-20 versus the sine of the angle θ in a two-slit interference experiment using light of wavelength 435 nm. The vertical axis scale is set by $\beta_s = 80.0 \text{ rad}$. What are (a) the slit separation, (b) the total number of interference maxima (count them on both sides of the pattern's center), (c) the smallest angle for a maxima, and (d) the greatest angle for a minimum? Assume that none of the interference maxima are completely eliminated by a diffraction minimum.

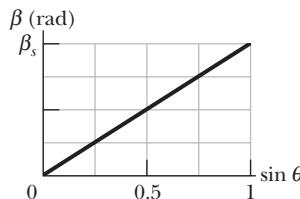


Fig. 36-45 Problem 40.

•41 In the two-slit interference experiment of Fig. 35-10, the slit widths are each $12.0 \mu\text{m}$, their separation is $24.0 \mu\text{m}$, the wavelength is 600 nm, and the viewing screen is at a distance of 4.00 m. Let I_p represent the intensity at point P on the screen, at height $y = 70.0 \text{ cm}$. (a) What is the ratio of I_p to the intensity I_m at the center of the pattern? (b) Determine where P is in the two-slit interference pattern by giving the maximum or minimum on which it lies or the maximum and minimum between which it lies. (c) In the same way, for the diffraction that occurs, determine where point P is in the diffraction pattern.

•42  (a) In a double-slit experiment, what ratio of d to a causes diffraction to eliminate the fourth bright side fringe? (b) What other bright fringes are also eliminated?

••43 SSM WWW (a) How many bright fringes appear between the first diffraction-envelope minima to either side of the central maximum in a double-slit pattern if $\lambda = 550 \text{ nm}$, $d = 0.150 \text{ mm}$, and $a = 30.0 \mu\text{m}$? (b) What is the ratio of the intensity of the third bright fringe to the intensity of the central fringe?

sec. 36-8 Diffraction Gratings

•44 Perhaps to confuse a predator, some tropical gyrid beetles (whirligig beetles) are colored by optical interference that is due to scales whose alignment forms a diffraction grating (which scatters light instead of transmitting it). When the incident light rays are perpendicular to the grating, the angle between the first-order maxima (on opposite sides of the zeroth-order maximum) is about 26° in light with a wavelength of 550 nm. What is the grating spacing of the beetle?

•45 A diffraction grating 20.0 mm wide has 6000 rulings. Light of wavelength 589 nm is incident perpendicularly on the grating. What are the (a) largest, (b) second largest, and (c) third largest values of θ at which maxima appear on a distant viewing screen?

•46 Visible light is incident perpendicularly on a grating with 315 rulings/mm. What is the longest wavelength that can be seen in the fifth-order diffraction?

•47 SSM ILW A grating has 400 lines/mm. How many orders of the entire visible spectrum (400–700 nm) can it produce in a diffraction experiment, in addition to the $m = 0$ order?

•48 A diffraction grating is made up of slits of width 300 nm with separation 900 nm. The grating is illuminated by monochromatic plane waves of wavelength $\lambda = 600 \text{ nm}$ at normal incidence. (a) How many maxima are there in the full diffraction pattern? (b) What is the angular width of a spectral line observed in the first order if the grating has 1000 slits?

•49 SSM WWW Light of wavelength 600 nm is incident normally on a diffraction grating. Two adjacent maxima occur at angles given by $\sin \theta = 0.2$ and $\sin \theta = 0.3$. The fourth-order maxima are missing. (a) What is the separation between adjacent slits? (b) What is the smallest slit width this grating can have? For that slit width, what are the (c) largest, (d) second largest, and (e) third largest values of the order number m of the maxima produced by the grating?

•50 With light from a gaseous discharge tube incident normally on a grating with slit separation $1.73 \mu\text{m}$, sharp maxima of green light are experimentally found at angles $\theta = \pm 17.6^\circ, 37.3^\circ, -37.1^\circ, 65.2^\circ$, and -65.0° . Compute the wavelength of the green light that best fits these data.

•51 A diffraction grating having 180 lines/mm is illuminated with a light signal containing only two wavelengths, $\lambda_1 = 400 \text{ nm}$ and $\lambda_2 = 500 \text{ nm}$. The signal is incident perpendicularly on the grating. (a) What is the angular separation between the second-order maxima of these two wavelengths? (b) What is the smallest angle at which two of the resulting maxima are superimposed? (c) What is the highest order for which maxima for both wavelengths are present in the diffraction pattern?

•52 GO A beam of light consisting of wavelengths from 460.0 nm to 640.0 nm is directed perpendicularly onto a diffraction grating with 160 lines/mm. (a) What is the lowest order that is overlapped by another order? (b) What is the highest order for which the complete wavelength range of the beam is present? In that highest order, at what angle does the light at wave-

length (c) 460.0 nm and (d) 640.0 nm appear? (e) What is the greatest angle at which the light at wavelength 460.0 nm appears?

•53 A grating has 350 rulings/mm and is illuminated at normal incidence by white light. A spectrum is formed on a screen 30.0 cm from the grating. If a hole 10.0 mm square is cut in the screen, its inner edge being 50.0 mm from the central maximum and parallel to it, what are the (a) shortest and (b) longest wavelengths of the light that passes through the hole?

•54 Derive this expression for the intensity pattern for a three-slit “grating”:

$$I = \frac{1}{9} I_m (1 + 4 \cos \phi + 4 \cos^2 \phi),$$

where $\phi = (2\pi d \sin \theta)/\lambda$ and $a \ll \lambda$.

sec. 36-9 Gratings: Dispersion and Resolving Power

•55 SSM ILW A source containing a mixture of hydrogen and deuterium atoms emits red light at two wavelengths whose mean is 656.3 nm and whose separation is 0.180 nm. Find the minimum number of lines needed in a diffraction grating that can resolve these lines in the first order.

•56 (a) How many rulings must a 4.00-cm-wide diffraction grating have to resolve the wavelengths 415.496 and 415.487 nm in the second order? (b) At what angle are the second-order maxima found?

•57 Light at wavelength 589 nm from a sodium lamp is incident perpendicularly on a grating with 40 000 rulings over width 76 nm. What are the first-order (a) dispersion D and (b) resolving power R , the second-order (c) D and (d) R , and the third-order (e) D and (f) R ?

•58 A grating has 600 rulings/mm and is 5.0 mm wide. (a) What is the smallest wavelength interval it can resolve in the third order at $\lambda = 500 \text{ nm}$? (b) How many higher orders of maxima can be seen?

•59 A diffraction grating with a width of 2.0 cm contains 1000 lines/cm across that width. For an incident wavelength of 600 nm, what is the smallest wavelength difference this grating can resolve in the second order?

•60 The D line in the spectrum of sodium is a doublet with wavelengths 589.0 and 589.6 nm. Calculate the minimum number of lines needed in a grating that will resolve this doublet in the second-order spectrum.

•61 With a particular grating the sodium doublet (589.00 nm and 589.59 nm) is viewed in the third order at 10° to the normal and is barely resolved. Find (a) the grating spacing and (b) the total width of the rulings.

•62 A diffraction grating illuminated by monochromatic light normal to the grating produces a certain line at angle θ . (a) What is the product of that line's half-width and the grating's resolving power? (b) Evaluate that product for the first order of a grating of slit separation 900 nm in light of wavelength 600 nm.

•63 Assume that the limits of the visible spectrum are arbitrarily chosen as 430 and 680 nm. Calculate the number of rulings per millimeter of a grating that will spread the first-order spectrum through an angle of 20.0° .

sec. 36-10 X-Ray Diffraction

•64 What is the smallest Bragg angle for x rays of wavelength 30 pm to reflect from reflecting planes spaced 0.30 nm apart in a calcite crystal?

1020 CHAPTER 36 DIFFRACTION

•65 An x-ray beam of wavelength A undergoes first-order reflection (Bragg law diffraction) from a crystal when its angle of incidence to a crystal face is 23° , and an x-ray beam of wavelength 97 pm undergoes third-order reflection when its angle of incidence to that face is 60° . Assuming that the two beams reflect from the same family of reflecting planes, find (a) the interplanar spacing and (b) the wavelength A .

•66 An x-ray beam of a certain wavelength is incident on a NaCl crystal, at 30.0° to a certain family of reflecting planes of spacing 39.8 pm. If the reflection from those planes is of the first order, what is the wavelength of the x rays?

•67 Figure 36-46 is a graph of intensity versus angular position θ for the diffraction of an x-ray beam by a crystal. The horizontal scale is set by $\theta_s = 2.00^\circ$. The beam consists of two wavelengths, and the spacing between the reflecting planes is 0.94 nm. What are the (a) shorter and (b) longer wavelengths in the beam?

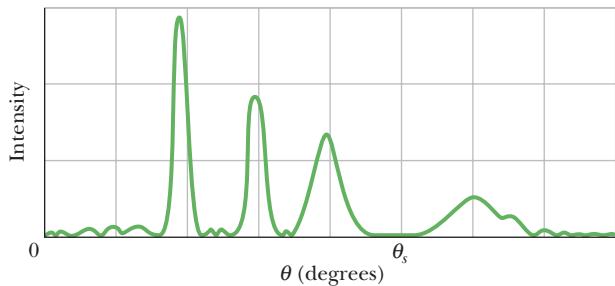


Fig. 36-46 Problem 67.

•68 If first-order reflection occurs in a crystal at Bragg angle 3.4° , at what Bragg angle does second-order reflection occur from the same family of reflecting planes?

•69 X rays of wavelength 0.12 nm are found to undergo second-order reflection at a Bragg angle of 28° from a lithium fluoride crystal. What is the interplanar spacing of the reflecting planes in the crystal?

•70 **GO** In Fig. 36-47, first-order reflection from the reflection planes shown occurs when an x-ray beam of wavelength 0.260 nm makes an angle $\theta = 63.8^\circ$ with the top face of the crystal. What is the unit cell size a_0 ?

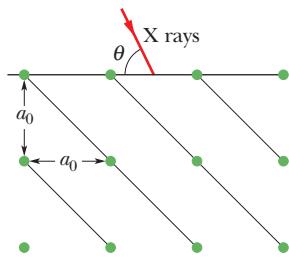


Fig. 36-47 Problem 70.

•71 **SSM** In Fig. 36-48, let a beam of x rays of wavelength 0.125 nm be incident on an NaCl crystal at angle $\theta = 45.0^\circ$ to the top face of the crystal and a family of reflecting planes. Let the reflecting planes have separation $d = 0.252$ nm. The crystal is turned through angle ϕ around an axis perpendicular to the plane of the

page until these reflecting planes give diffraction maxima. What are the (a) smaller and (b) larger value of ϕ if the crystal is turned clockwise and the (c) smaller and (d) larger value of ϕ if it is turned counterclockwise?

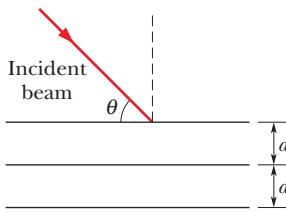


Fig. 36-48 Problems 71 and 72.

•72 In Fig. 36-48, an x-ray beam of wavelengths from 95.0 to 140 pm is incident at $\theta = 45.0^\circ$ to a family of reflecting planes with spacing $d = 275$ pm. What are the (a) longest wavelength λ and (b) associated order number m and the (c) shortest λ and (d) associated m of the intensity maxima in the diffraction of the beam?

•73 Consider a two-dimensional square crystal structure, such as one side of the structure shown in Fig. 36-28a. The largest interplanar spacing of reflecting planes is the unit cell size a_0 . Calculate and sketch the (a) second largest, (b) third largest, (c) fourth largest, (d) fifth largest, and (e) sixth largest interplanar spacing. (f) Show that your results in (a) through (e) are consistent with the general formula

$$d = \frac{a_0}{\sqrt{h^2 + k^2}},$$

where h and k are relatively prime integers (they have no common factor other than unity).

Additional Problems

74 An astronaut in a space shuttle claims she can just barely resolve two point sources on Earth's surface, 160 km below. Calculate their (a) angular and (b) linear separation, assuming ideal conditions. Take $\lambda = 540$ nm and the pupil diameter of the astronaut's eye to be 5.0 mm.

75 **SSM** Visible light is incident perpendicularly on a diffraction grating of 200 rulings/mm. What are the (a) longest, (b) second longest, and (c) third longest wavelengths that can be associated with an intensity maximum at $\theta = 30.0^\circ$?

76 A beam of light consists of two wavelengths, 590.159 nm and 590.220 nm, that are to be resolved with a diffraction grating. If the grating has lines across a width of 3.80 cm, what is the minimum number of lines required for the two wavelengths to be resolved in the second order?

77 **SSM** In a single-slit diffraction experiment, there is a minimum of intensity for orange light ($\lambda = 600$ nm) and a minimum of intensity for blue-green light ($\lambda = 500$ nm) at the same angle of 1.00 mrad. For what minimum slit width is this possible?

78 **GO** A double-slit system with individual slit widths of 0.030 mm and a slit separation of 0.18 mm is illuminated with 500 nm light directed perpendicular to the plane of the slits. What is the total number of complete bright fringes appearing between the two first-order minima of the diffraction pattern? (Do not count the fringes that coincide with the minima of the diffraction pattern.)

79 **SSM** A diffraction grating has resolving power $R = \lambda_{\text{avg}}/\Delta\lambda = Nm$. (a) Show that the corresponding frequency range Δf that can just be resolved is given by $\Delta f = c/Nm\lambda$. (b) From Fig. 36-22, show that the times required for light to travel along the ray at the bottom of the figure and the ray at the top differ by $\Delta t = (Nd/c) \sin \theta$. (c) Show that $(\Delta f)(\Delta t) = 1$, this relation being independent of the various grating parameters. Assume $N \gg 1$.

80 The pupil of a person's eye has a diameter of 5.00 mm. According to Rayleigh's criterion, what distance apart must two small objects be if their images are just barely resolved when they are 250 mm from the eye? Assume they are illuminated with light of wavelength 500 nm.

81 Light is incident on a grating at an angle ψ as shown in Fig. 36-49. Show that bright fringes occur at angles θ that satisfy the equation

$$d(\sin \psi + \sin \theta) = m\lambda, \quad \text{for } m = 0, 1, 2, \dots$$

(Compare this equation with Eq. 36-25.) Only the special case $\psi = 0$ has been treated in this chapter.

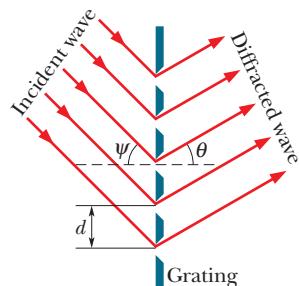


Fig. 36-49 Problem 81.

82 A grating with $d = 1.50 \mu\text{m}$ is illuminated at various angles of incidence by light of wavelength 600 nm. Plot, as a function of the angle of incidence (0 to 90°), the angular deviation of the first-order maximum from the incident direction. (See Problem 81.)

83 SSM In two-slit interference, if the slit separation is $14 \mu\text{m}$ and the slit widths are each $2.0 \mu\text{m}$, (a) how many two-slit maxima are in the central peak of the diffraction envelope and (b) how many are in either of the first side peak of the diffraction envelope?

84 In a two-slit interference pattern, what is the ratio of slit separation to slit width if there are 17 bright fringes within the central diffraction envelope and the diffraction minima coincide with two-slit interference maxima?

85 A beam of light with a narrow wavelength range centered on 450 nm is incident perpendicularly on a diffraction grating with a width of 1.80 cm and a line density of 1400 lines/cm across that width. For this light, what is the smallest wavelength difference this grating can resolve in the third order?

86 If you look at something 40 m from you, what is the smallest length (perpendicular to your line of sight) that you can resolve, according to Rayleigh's criterion? Assume the pupil of your eye has a diameter of 4.00 mm, and use 500 nm as the wavelength of the light reaching you.

87 Two yellow flowers are separated by 60 cm along a line perpendicular to your line of sight to the flowers. How far are you from the flowers when they are at the limit of resolution according to the Rayleigh criterion? Assume the light from the flowers has a single wavelength of 550 nm and that your pupil has a diameter of 5.5 mm.

88 In a single-slit diffraction experiment, what must be the ratio of the slit width to the wavelength if the second diffraction minima are to occur at an angle of 37.0° from the center of the diffraction pattern on a viewing screen?

89 A diffraction grating 3.00 cm wide produces the second order at 33.0° with light of wavelength 600 nm. What is the total number of lines on the grating?

90 A single-slit diffraction experiment is set up with light of wavelength 420 nm, incident perpendicularly on a slit of width $5.10 \mu\text{m}$. The viewing screen is 3.20 m distant. On the screen, what is the distance between the center of the diffraction pattern and the second diffraction minimum?

91 A diffraction grating has 8900 slits across 1.20 cm. If light with a wavelength of 500 nm is sent through it, how many orders (maxima) lie to one side of the central maximum?

92 In an experiment to monitor the Moon's surface with a light beam, pulsed radiation from a ruby laser ($\lambda = 0.69 \mu\text{m}$) was directed to the Moon through a reflecting telescope with a mirror radius of 1.3 m. A reflector on the Moon behaved like a circular flat mirror with radius 10 cm, reflecting the light directly back toward the telescope on Earth. The reflected light was then detected after being brought to a focus by this telescope. Approximately what fraction of the original light energy was picked up by the detector? Assume that for each direction of travel all the energy is in the central diffraction peak.

93 In June 1985, a laser beam was sent out from the Air Force Optical Station on Maui, Hawaii, and reflected back from the shuttle *Discovery* as it sped by 354 km overhead. The diameter of the central maximum of the beam at the shuttle position was said to be 9.1 m, and the beam wavelength was 500 nm. What is the effective diameter of the laser aperture at the Maui ground station? (Hint: A laser beam spreads only because of diffraction; assume a circular exit aperture.)

94 A diffraction grating 1.00 cm wide has 10 000 parallel slits. Monochromatic light that is incident normally is diffracted through 30° in the first order. What is the wavelength of the light?

95 SSM If you double the width of a single slit, the intensity of the central maximum of the diffraction pattern increases by a factor of 4, even though the energy passing through the slit only doubles. Explain this quantitatively.

96 When monochromatic light is incident on a slit $22.0 \mu\text{m}$ wide, the first diffraction minimum lies at 1.80° from the direction of the incident light. What is the wavelength?

97 A spy satellite orbiting at 160 km above Earth's surface has a lens with a focal length of 3.6 m and can resolve objects on the ground as small as 30 cm. For example, it can easily measure the size of an aircraft's air intake port. What is the effective diameter of the lens as determined by diffraction consideration alone? Assume $\lambda = 550 \text{ nm}$.

98 Suppose that two points are separated by 2.0 cm. If they are viewed by an eye with a pupil opening of 5.0 mm, what distance from the viewer puts them at the Rayleigh limit of resolution? Assume a light wavelength of 500 nm.

99 A diffraction grating has 200 lines/mm. Light consisting of a continuous range of wavelengths between 550 nm and 700 nm is incident perpendicularly on the grating. (a) What is the lowest order that is overlapped by another order? (b) What is the highest order for which the complete spectrum is present?

37

RELATIVITY

37-1 WHAT IS PHYSICS?

One principal subject of physics is **relativity**, the field of study that measures events (things that happen): where and when they happen, and by how much any two events are separated in space and in time. In addition, relativity has to do with transforming such measurements (and also measurements of energy and momentum) between reference frames that move relative to each other. (Hence the name *relativity*.)

Transformations and moving reference frames, such as those we discussed in Sections 4-8 and 4-9, were well understood and quite routine to physicists in 1905. Then Albert Einstein (Fig. 37-1) published his **special theory of relativity**. The adjective *special* means that the theory deals only with **inertial reference frames**, which are frames in which Newton's laws are valid. (Einstein's *general theory of relativity* treats the more challenging situation in which reference frames can undergo gravitational acceleration; in this chapter the term *relativity* implies only inertial reference frames.)

Starting with two deceptively simple postulates, Einstein stunned the scientific world by showing that the old ideas about relativity were wrong, even though everyone was so accustomed to them that they seemed to be unquestionable common sense. This supposed common sense, however, was derived only from experience with things that move rather slowly. Einstein's relativity, which turns out to be correct for all physically possible speeds, predicted many effects that were, at first study, bizarre because no one had ever experienced them.

In particular, Einstein demonstrated that space and time are entangled; that is, the time between two events depends on how far apart they occur, and vice versa. Also, the entanglement is different for observers who move relative to each other. One result is that time does not pass at a fixed rate, as if it were ticked off with mechanical regularity on some master grandfather clock that controls the universe. Rather, that rate is adjustable: Relative motion can change the rate at which time passes. Prior to 1905, no one but a few daydreamers would have thought that. Now, engineers and scientists take it for granted because their experience with special relativity has reshaped their common sense. For example, any engineer involved with the Global Positioning System of the NAVSTAR satellites must routinely use relativity (both special relativity and general relativity) to determine the rate at which time passes on the satellites because that rate differs from the rate on Earth's surface. If the engineers failed to take relativity into account, GPS would become almost useless in less than one day.

Special relativity has the reputation of being difficult. It is not difficult mathematically, at least not here. However, it is difficult in that we must be very careful about *who* measures *what* about an event and just *how* that measurement is made—and it can be difficult because it can contradict routine experience.



Fig. 37-1 Einstein posing for a photograph as fame began to accumulate. (Corbis Images)

37-2 The Postulates

We now examine the two postulates of relativity, on which Einstein's theory is based:



- 1. The Relativity Postulate:** The laws of physics are the same for observers in all inertial reference frames. No one frame is preferred over any other.

Galileo assumed that the laws of *mechanics* were the same in all inertial reference frames. Einstein extended that idea to include *all* the laws of physics, especially those of electromagnetism and optics. This postulate does *not* say that the measured values of all physical quantities are the same for all inertial observers; most are not the same. It is the *laws of physics*, which relate these measurements to one another, that are the same.



- 2. The Speed of Light Postulate:** The speed of light in vacuum has the same value c in all directions and in all inertial reference frames.

We can also phrase this postulate to say that there is in nature an *ultimate speed* c , the same in all directions and in all inertial reference frames. Light happens to travel at this ultimate speed. However, no entity that carries energy or information can exceed this limit. Moreover, no particle that has mass can actually reach speed c , no matter how much or for how long that particle is accelerated. (Alas, the faster-than-light warp drive used in many science fiction stories appears to be impossible.)

Both postulates have been exhaustively tested, and no exceptions have ever been found.

The Ultimate Speed

The existence of a limit to the speed of accelerated electrons was shown in a 1964 experiment by W. Bertozzi, who accelerated electrons to various measured speeds and—by an independent method—measured their kinetic energies. He found that as the force on a very fast electron is increased, the electron's measured kinetic energy increases toward very large values but its speed does not increase appreciably (Fig. 37-2). Electrons have been accelerated in laboratories to at least 0.999 999 999 95 times the speed of light but—close though it may be—that speed is still less than the ultimate speed c .

This ultimate speed has been defined to be exactly

$$c = 299\,792\,458 \text{ m/s.} \quad (37-1)$$

Caution: So far in this book we have (appropriately) approximated c as 3.0×10^8 m/s, but in this chapter we shall often use the exact value. You might want to store the exact value in your calculator's memory (if it is not there already), to be called up when needed.

Testing the Speed of Light Postulate

If the speed of light is the same in all inertial reference frames, then the speed of light emitted by a source relative to, say, a laboratory should be the same as the speed of light that is emitted by a source at rest in the laboratory. This claim has been tested directly, in an experiment of high precision. The “light source” was the *neutral pion* (symbol π^0), an unstable, short-lived particle that can be produced by collisions in a particle accelerator. It decays (transforms) into two gamma rays by the process



Gamma rays are part of the electromagnetic spectrum (at very high frequencies)

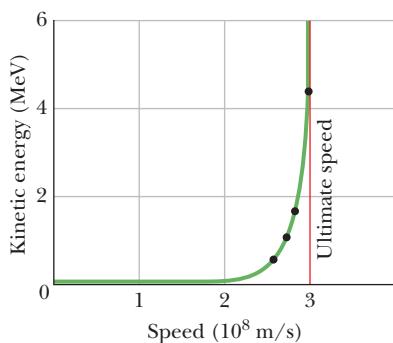


Fig. 37-2 The dots show measured values of the kinetic energy of an electron plotted against its measured speed. No matter how much energy is given to an electron (or to any other particle having mass), its speed can never equal or exceed the ultimate limiting speed c . (The plotted curve through the dots shows the predictions of Einstein's special theory of relativity.)

and so obey the speed of light postulate, just as visible light does. (In this chapter we shall use the term light for any type of electromagnetic wave, visible or not.)

In 1964, physicists at CERN, the European particle-physics laboratory near Geneva, generated a beam of pions moving at a speed of $0.999\ 75c$ with respect to the laboratory. The experimenters then measured the speed of the gamma rays emitted from these very rapidly moving sources. They found that the speed of the light emitted by the pions was the same as it would be if the pions were at rest in the laboratory, namely c .

37-3 Measuring an Event

An **event** is something that happens, and every event can be assigned three space coordinates and one time coordinate. Among many possible events are (1) the turning on or off of a tiny lightbulb, (2) the collision of two particles, (3) the passage of a pulse of light through a specified point, (4) an explosion, and (5) the sweeping of the hand of a clock past a marker on the rim of the clock. A certain observer, fixed in a certain inertial reference frame, might, for example, assign to an event *A* the coordinates given in Table 37-1. Because space and time are entangled with each other in relativity, we can describe these coordinates collectively as *spacetime* coordinates. The coordinate system itself is part of the reference frame of the observer.

A given event may be recorded by any number of observers, each in a different inertial reference frame. In general, different observers will assign different spacetime coordinates to the same event. Note that an event does not “belong” to any particular inertial reference frame. An event is just something that happens, and anyone in any reference frame may detect it and assign spacetime coordinates to it.

Making such an assignment can be complicated by a practical problem. For example, suppose a balloon bursts 1 km to your right while a firecracker pops 2 km to your left, both at 9:00 A.M. However, you do not detect either event precisely at 9:00 A.M. because at that instant light from the events has not yet reached you. Because light from the firecracker pop has farther to go, it arrives at your eyes later than does light from the balloon burst, and thus the pop will seem to have occurred later than the burst. To sort out the actual times and to assign 9:00 A.M. as the happening time for both events, you must calculate the travel times of the light and then subtract these times from the arrival times.

This procedure can be very messy in more challenging situations, and we need an easier procedure that automatically eliminates any concern about the travel time from an event to an observer. To set up such a procedure, we shall construct an imaginary array of measuring rods and clocks throughout the observer’s inertial frame (the array moves rigidly with the observer). This construction may seem contrived, but it spares us much confusion and calculation and allows us to find the coordinates, as follows.

1. The Space Coordinates. We imagine the observer’s coordinate system fitted with a close-packed, three-dimensional array of measuring rods, one set of rods parallel to each of the three coordinate axes. These rods provide a way to determine coordinates along the axes. Thus, if the event is, say, the turning on of a small lightbulb, the observer, in order to locate the position of the event, need only read the three space coordinates at the bulb’s location.

2. The Time Coordinate. For the time coordinate, we imagine that every point of intersection in the array of measuring rods includes a tiny clock, which the observer can read because the clock is illuminated by the light generated by the event. Figure 37-3 suggests one plane in the “jungle gym” of clocks and measuring rods we have described.

Table 37-1

Record of Event A

Coordinate	Value
<i>x</i>	3.58 m
<i>y</i>	1.29 m
<i>z</i>	0 m
<i>t</i>	34.5 s

The array of clocks must be synchronized properly. It is not enough to assemble a set of identical clocks, set them all to the same time, and then move them to their assigned positions. We do not know, for example, whether moving the clocks will change their rates. (Actually, it will.) We must put the clocks in place and *then* synchronize them.

If we had a method of transmitting signals at infinite speed, synchronization would be a simple matter. However, no known signal has this property. We therefore choose light (any part of the electromagnetic spectrum) to send out our synchronizing signals because, in vacuum, light travels at the greatest possible speed, the limiting speed c .

Here is one of many ways in which an observer might synchronize an array of clocks using light signals: The observer enlists the help of a great number of temporary helpers, one for each clock. The observer then stands at a point selected as the origin and sends out a pulse of light when the origin clock reads $t = 0$. When the light pulse reaches the location of a helper, that helper sets the clock there to read $t = r/c$, where r is the distance between the helper and the origin. The clocks are then synchronized.

- 3. The Spacetime Coordinates.** The observer can now assign spacetime coordinates to an event by simply recording the time on the clock nearest the event and the position as measured on the nearest measuring rods. If there are two events, the observer computes their separation in time as the difference in the times on clocks near each and their separation in space from the differences in coordinates on rods near each. We thus avoid the practical problem of calculating the travel times of the signals to the observer from the events.

37-4 The Relativity of Simultaneity

Suppose that one observer (Sam) notes that two independent events (event Red and event Blue) occur at the same time. Suppose also that another observer (Sally), who is moving at a constant velocity \vec{v} with respect to Sam, also records these same two events. Will Sally also find that they occur at the same time?

The answer is that in general she will not:

 If two observers are in relative motion, they will not, in general, agree as to whether two events are simultaneous. If one observer finds them to be simultaneous, the other generally will not.

We cannot say that one observer is right and the other wrong. Their observations are equally valid, and there is no reason to favor one over the other.

The realization that two contradictory statements about the same natural event can be correct is a seemingly strange outcome of Einstein's theory. However, in Chapter 17 we saw another way in which motion can affect measurement without balking at the contradictory results: In the Doppler effect, the frequency an observer measures for a sound wave depends on the relative motion of observer and source. Thus, two observers moving relative to each other can measure different frequencies for the same wave, and both measurements are correct.

We conclude the following:

 Simultaneity is not an absolute concept but rather a relative one, depending on the motion of the observer.

If the relative speed of the observers is very much less than the speed of light, then measured departures from simultaneity are so small that they are not noticeable. Such is the case for all our experiences of daily living; that is why the relativity of simultaneity is unfamiliar.

We use this array to assign spacetime coordinates.

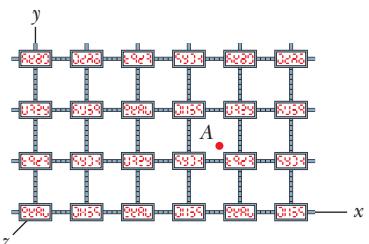


Fig. 37-3 One section of a three-dimensional array of clocks and measuring rods by which an observer can assign space-time coordinates to an event, such as a flash of light at point A . The event's space coordinates are approximately $x = 3.6$ rod lengths, $y = 1.3$ rod lengths, and $z = 0$. The time coordinate is whatever time appears on the clock closest to A at the instant of the flash.

A Closer Look at Simultaneity

Let us clarify the relativity of simultaneity with an example based on the postulates of relativity, no clocks or measuring rods being directly involved. Figure 37-4 shows two long spaceships (the SS *Sally* and the SS *Sam*), which can serve as inertial reference frames for observers Sally and Sam. The two observers are stationed at the midpoints of their ships. The ships are separating along a common *x* axis, the relative velocity of *Sally* with respect to *Sam* being \vec{v} . Figure 37-4a shows the ships with the two observer stations momentarily aligned opposite each other.

Two large meteorites strike the ships, one setting off a red flare (event Red) and the other a blue flare (event Blue), not necessarily simultaneously. Each event leaves a permanent mark on each ship, at positions RR' and BB' .

Let us suppose that the expanding wavefronts from the two events happen to reach Sam at the same time, as Fig. 37-4b shows. Let us further suppose that, after the episode, Sam finds, by measuring the marks on his spaceship, that he was indeed stationed exactly halfway between the markers B and R on his ship when the two events occurred. He will say:

Sam Light from event Red and light from event Blue reached me at the same time.
From the marks on my spaceship, I find that I was standing halfway between the two sources. Therefore, event Red and event Blue were simultaneous events.

As study of Fig. 37-4 shows, Sally and the expanding wavefront from event Red are moving *toward* each other, while she and the expanding wavefront from event Blue are moving in the *same direction*. Thus, the wavefront from event Red will reach Sally *before* the wavefront from event Blue does. She will say:

Sally Light from event Red reached me before light from event Blue did. From the marks on my spaceship, I found that I too was standing halfway between the two sources. Therefore, the events were not simultaneous; event Red occurred first, followed by event Blue.

These reports do not agree. Nevertheless, *both* observers are correct.

Note carefully that there is only one wavefront expanding from the site of each event and that *this wavefront travels with the same speed c in both reference frames*, exactly as the speed of light postulate requires.

It *might* have happened that the meteorites struck the ships in such a way that the two hits appeared to Sally to be simultaneous. If that had been the case, then Sam would have declared them not to be simultaneous.

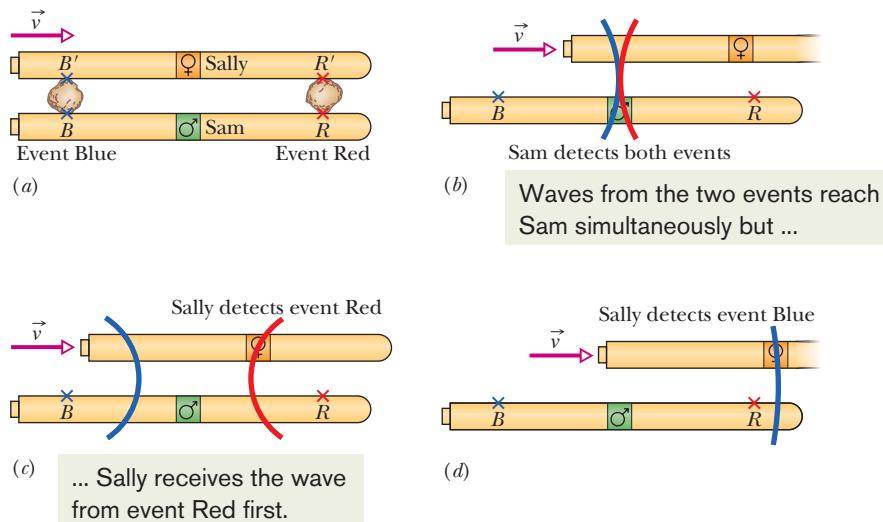


Fig. 37-4 The spaceships of Sally and Sam and the occurrences of events from Sam's view. Sally's ship moves rightward with velocity \vec{v} . (a) Event Red occurs at positions RR' and event Blue occurs at positions BB' ; each event sends out a wave of light. (b) Sam simultaneously detects the waves from event Red and event Blue. (c) Sally detects the wave from event Red. (d) Sally detects the wave from event Blue.

37-5 The Relativity of Time

If observers who move relative to each other measure the time interval (or *temporal separation*) between two events, they generally will find different results. Why? Because the spatial separation of the events can affect the time intervals measured by the observers.



The time interval between two events depends on how far apart they occur in both space and time; that is, their spatial and temporal separations are entangled.

In this section we discuss this entanglement by means of an example; however, the example is restricted in a crucial way: *To one of two observers, the two events occur at the same location.* We shall not get to more general examples until Section 37-7.

Figure 37-5a shows the basics of an experiment Sally conducts while she and her equipment—a light source, a mirror, and a clock—ride in a train moving with constant velocity v relative to a station. A pulse of light leaves the light source B (event 1), travels vertically upward, is reflected vertically downward by the mirror, and then is detected back at the source (event 2). Sally measures a certain time interval Δt_0 between the two events, related to the distance D from source to mirror by

$$\Delta t_0 = \frac{2D}{c} \quad (\text{Sally}). \quad (37-3)$$

The two events occur at the same location in Sally's reference frame, and she needs only one clock C at that location to measure the time interval. Clock C is shown twice in Fig. 37-5a, at the beginning and end of the interval.

Consider now how these same two events are measured by Sam, who is standing on the station platform as the train passes. Because the equipment moves with the train during the travel time of the light, Sam sees the path of the light as shown in Fig. 37-5b. For him, the two events occur at different places in his reference frame, and so to measure the time interval between events, Sam must use *two* synchronized clocks, C_1 and C_2 , one at each event. According to Einstein's speed of light postulate, the light travels at the same speed c for Sam as

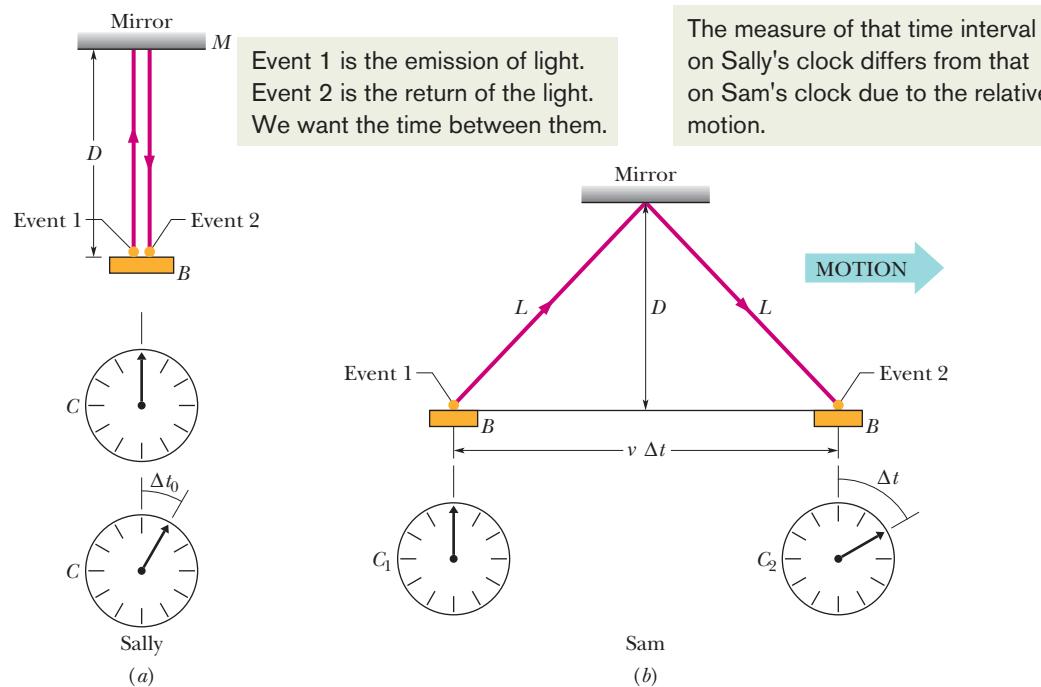


Fig. 37-5 (a) Sally, on the train, measures the time interval Δt_0 between events 1 and 2 using a single clock C on the train. That clock is shown twice: first for event 1 and then for event 2. (b) Sam, watching from the station as the events occur, requires two synchronized clocks, C_1 at event 1 and C_2 at event 2, to measure the time interval between the two events; his measured time interval is Δt .

for Sally. Now, however, the light travels distance $2L$ between events 1 and 2. The time interval measured by Sam between the two events is

$$\Delta t = \frac{2L}{c} \quad (\text{Sam}), \quad (37-4)$$

in which $L = \sqrt{(\frac{1}{2}v \Delta t)^2 + D^2}$. (37-5)

From Eq. 37-3, we can write this as

$$L = \sqrt{(\frac{1}{2}v \Delta t)^2 + (\frac{1}{2}c \Delta t_0)^2}. \quad (37-6)$$

If we eliminate L between Eqs. 37-4 and 37-6 and solve for Δt , we find

$$\Delta t = \frac{\Delta t_0}{\sqrt{1 - (v/c)^2}}. \quad (37-7)$$

Equation 37-7 tells us how Sam's measured interval Δt between the events compares with Sally's interval Δt_0 . Because v must be less than c , the denominator in Eq. 37-7 must be less than unity. Thus, Δt must be greater than Δt_0 : Sam measures a *greater* time interval between the two events than does Sally. Sam and Sally have measured the time interval between the *same* two events, but the relative motion between Sam and Sally made their measurements *different*. We conclude that relative motion can change the *rate* at which time passes between two events; the key to this effect is the fact that the speed of light is the same for both observers.

We distinguish between the measurements of Sam and Sally in this way:



When two events occur at the same location in an inertial reference frame, the time interval between them, measured in that frame, is called the **proper time interval** or the **proper time**. Measurements of the same time interval from any other inertial reference frame are always greater.

Thus, Sally measures a proper time interval, and Sam measures a greater time interval. (The term *proper* is unfortunate in that it implies that any other measurement is improper or nonreal. That is just not so.) The amount by which a measured time interval is greater than the corresponding proper time interval is called **time dilation**. (To dilate is to expand or stretch; here the time interval is expanded or stretched.)

Often the dimensionless ratio v/c in Eq. 37-7 is replaced with β , called the **speed parameter**, and the dimensionless inverse square root in Eq. 37-7 is often replaced with γ , called the **Lorentz factor**:

$$\gamma = \frac{1}{\sqrt{1 - \beta^2}} = \frac{1}{\sqrt{1 - (v/c)^2}}. \quad (37-8)$$

With these replacements, we can rewrite Eq. 37-7 as

$$\Delta t = \gamma \Delta t_0 \quad (\text{time dilation}). \quad (37-9)$$

The speed parameter β is always less than unity, and, provided v is not zero, γ is always greater than unity. However, the difference between γ and 1 is not significant unless $v > 0.1c$. Thus, in general, “old relativity” works well enough for $v < 0.1c$, but we must use special relativity for greater values of v . As shown in Fig. 37-6, γ increases rapidly in magnitude as β approaches 1 (as v approaches c). Therefore, the greater the relative speed between Sally and Sam is, the greater will be the time interval measured by Sam, until at a great enough speed, the interval takes “forever.”

As the speed parameter goes to 1.0 (as the speed approaches c), the Lorentz factor approaches infinity.

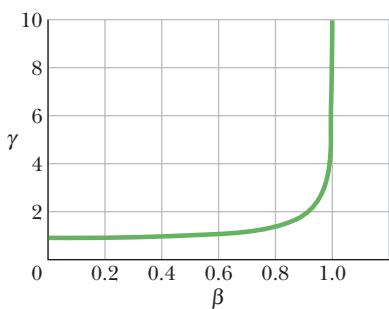


Fig. 37-6 A plot of the Lorentz factor γ as a function of the speed parameter $\beta (= v/c)$.

You might wonder what Sally says about Sam's having measured a greater time interval than she did. His measurement comes as no surprise to her, because to her, he failed to synchronize his clocks C_1 and C_2 in spite of his insistence that he did. Recall that observers in relative motion generally do not agree about simultaneity. Here, Sam insists that his two clocks simultaneously read the same time when event 1 occurred. To Sally, however, Sam's clock C_2 was erroneously set ahead during the synchronization process. Thus, when Sam read the time of event 2 on it, to Sally he was reading off a time that was too large, and that is why the time interval he measured between the two events was greater than the interval she measured.

Two Tests of Time Dilation

- 1. Microscopic Clocks.** Subatomic particles called *muons* are unstable; that is, when a muon is produced, it lasts for only a short time before it *decays* (transforms into particles of other types). The *lifetime* of a muon is the time interval between its production (event 1) and its decay (event 2). When muons are stationary and their lifetimes are measured with stationary clocks (say, in a laboratory), their average lifetime is $2.200 \mu\text{s}$. This is a proper time interval because, for each muon, events 1 and 2 occur at the same location in the reference frame of the muon—namely, at the muon itself. We can represent this proper time interval with Δt_0 ; moreover, we can call the reference frame in which it is measured the *rest frame* of the muon.

If, instead, the muons are moving, say, through a laboratory, then measurements of their lifetimes made with the laboratory clocks should yield a greater average lifetime (a dilated average lifetime). To check this conclusion, measurements were made of the average lifetime of muons moving with a speed of $0.9994c$ relative to laboratory clocks. From Eq. 37-8, with $\beta = 0.9994$, the Lorentz factor for this speed is

$$\gamma = \frac{1}{\sqrt{1 - \beta^2}} = \frac{1}{\sqrt{1 - (0.9994)^2}} = 28.87.$$

Equation 37-9 then yields, for the average dilated lifetime,

$$\Delta t = \gamma \Delta t_0 = (28.87)(2.200 \mu\text{s}) = 63.51 \mu\text{s}.$$

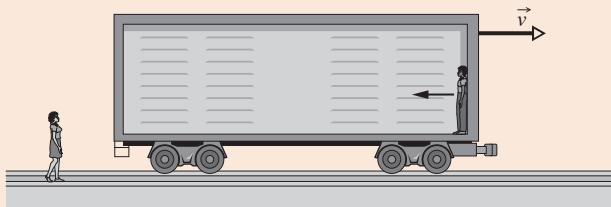
The actual measured value matched this result within experimental error.

- 2. Macroscopic Clocks.** In October 1971, Joseph Hafele and Richard Keating carried out what must have been a grueling experiment. They flew four portable atomic clocks twice around the world on commercial airlines, once in each direction. Their purpose was “to test Einstein’s theory of relativity with macroscopic clocks.” As we have just seen, the time dilation predictions of Einstein’s theory have been confirmed on a microscopic scale, but there is great comfort in seeing a confirmation made with an actual clock. Such macroscopic measurements became possible only because of the very high precision of modern atomic clocks. Hafele and Keating verified the predictions of the theory to within 10%. (Einstein’s *general* theory of relativity, which predicts that the rate at which time passes on a clock is influenced by the gravitational force on the clock, also plays a role in this experiment.)

A few years later, physicists at the University of Maryland carried out a similar experiment with improved precision. They flew an atomic clock round and round over Chesapeake Bay for flights lasting 15 h and succeeded in checking the time dilation prediction to better than 1%. Today, when atomic clocks are transported from one place to another for calibration or other purposes, the time dilation caused by their motion is always taken into account.

CHECKPOINT 1

Standing beside railroad tracks, we are suddenly startled by a relativistic boxcar traveling past us as shown in the figure. Inside, a well-equipped hobo fires a laser pulse from the front of the boxcar to its rear. (a) Is our measurement of the speed of the pulse greater than, less than, or the same as that measured by the hobo? (b) Is his measurement of the flight time of the pulse a proper time? (c) Are his measurement and our measurement of the flight time related by Eq. 37-9?

**Sample Problem****Time dilation for a space traveler who returns to Earth**

Your starship passes Earth with a relative speed of $0.9990c$. After traveling 10.0 y (your time), you stop at lookout post LP13, turn, and then travel back to Earth with the same relative speed. The trip back takes another 10.0 y (your time). How long does the round trip take according to measurements made on Earth? (Neglect any effects due to the accelerations involved with stopping, turning, and getting back up to speed.)

KEY IDEAS

We begin by analyzing the outward trip:

1. This problem involves measurements made from two (inertial) reference frames, one attached to Earth and the other (your reference frame) attached to your ship.
2. The outward trip involves two events: the start of the trip at Earth and the end of the trip at LP13.
3. Your measurement of 10.0 y for the outward trip is the proper time Δt_0 between those two events, because the events occur at the same location in your reference frame—namely, on your ship.

4. The Earth-frame measurement of the time interval Δt for the outward trip must be greater than Δt_0 , according to Eq. 37-9 ($\Delta t = \gamma \Delta t_0$) for time dilation.

Calculations: Using Eq. 37-8 to substitute for γ in Eq. 37-9, we find

$$\begin{aligned}\Delta t &= \frac{\Delta t_0}{\sqrt{1 - (v/c)^2}} \\ &= \frac{10.0\text{ y}}{\sqrt{1 - (0.9990c/c)^2}} = (22.37)(10.0\text{ y}) = 224\text{ y}.\end{aligned}$$

On the return trip, we have the same situation and the same data. Thus, the round trip requires 20 y of your time but

$$\Delta t_{\text{total}} = (2)(224\text{ y}) = 448\text{ y} \quad (\text{Answer})$$

of Earth time. In other words, you have aged 20 y while the Earth has aged 448 y . Although you cannot travel into the past (as far as we know), you can travel into the future of, say, Earth, by using high-speed relative motion to adjust the rate at which time passes.

Sample Problem**Time dilation and travel distance for a relativistic particle**

The elementary particle known as the *positive kaon* (K^+) is unstable in that it can *decay* (transform) into other particles. Although the decay occurs randomly, we find that, on average, a positive kaon has a lifetime of $0.1237\text{ }\mu\text{s}$ when stationary—that is, when the lifetime is measured in the rest frame of the kaon. If a positive kaon has a speed of $0.990c$ relative to a laboratory reference frame when the kaon is produced, how far can it travel in that frame during its lifetime according to *classical physics* (which is a reasonable approximation for speeds much less than c) and according to special relativity (which is correct for all physically possible speeds)?

KEY IDEAS

1. We have two (inertial) reference frames, one attached to the kaon and the other attached to the laboratory.
2. This problem also involves two events: the start of the kaon's travel (when the kaon is produced) and the end of that travel (at the end of the kaon's lifetime).
3. The distance traveled by the kaon between those two events is related to its speed v and the time interval for the travel by

$$v = \frac{\text{distance}}{\text{time interval}}. \quad (37-10)$$

With these ideas in mind, let us solve for the distance first with classical physics and then with special relativity.

Classical physics: In classical physics we would find the same distance and time interval (in Eq. 37-10) whether we measured them from the kaon frame or from the laboratory frame. Thus, we need not be careful about the frame in which the measurements are made. To find the kaon's travel distance d_{cp} according to classical physics, we first rewrite Eq. 37-10 as

$$d_{\text{cp}} = v \Delta t, \quad (37-11)$$

where Δt is the time interval between the two events in either frame. Then, substituting 0.990c for v and 0.1237 μs for Δt in Eq. 37-11, we find

$$\begin{aligned} d_{\text{cp}} &= (0.990c) \Delta t \\ &= (0.990)(299\,792\,458 \text{ m/s})(0.1237 \times 10^{-6} \text{ s}) \\ &= 36.7 \text{ m.} \end{aligned} \quad (\text{Answer})$$

This is how far the kaon would travel if classical physics were correct at speeds close to c .

Special relativity: In special relativity we must be very careful that both the distance and the time interval in Eq. 37-10 are measured in the *same* reference frame—especially when the speed is close to c , as here. Thus, to find the actual travel distance d_{sr} of the kaon *as measured from the laboratory frame* and according to special relativity, we rewrite Eq. 37-10 as

$$d_{\text{sr}} = v \Delta t, \quad (37-12)$$



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where Δt is the time interval between the two events *as measured from the laboratory frame*.

Before we can evaluate d_{sr} in Eq. 37-12, we must find Δt . The 0.1237 μs time interval is a proper time because the two events occur at the same location in the kaon frame—namely, at the kaon itself. Therefore, let Δt_0 represent this proper time interval. Then we can use Eq. 37-9 ($\Delta t = \gamma \Delta t_0$) for time dilation to find the time interval Δt as measured from the laboratory frame. Using Eq. 37-8 to substitute for γ in Eq. 37-9 leads to

$$\Delta t = \frac{\Delta t_0}{\sqrt{1 - (v/c)^2}} = \frac{0.1237 \times 10^{-6} \text{ s}}{\sqrt{1 - (0.990c/c)^2}} = 8.769 \times 10^{-7} \text{ s.}$$

This is about seven times longer than the kaon's proper lifetime. That is, the kaon's lifetime is about seven times longer in the laboratory frame than in its own frame—the kaon's lifetime is dilated. We can now evaluate Eq. 37-12 for the travel distance d_{sr} in the laboratory frame as

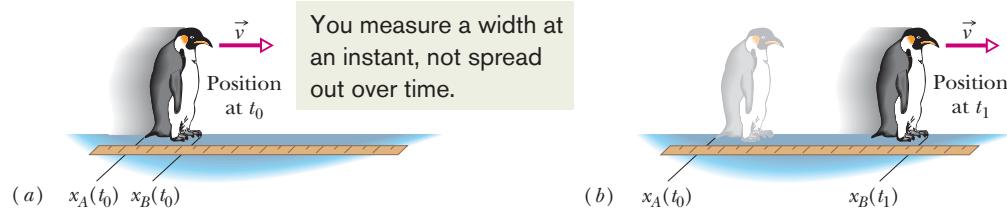
$$\begin{aligned} d_{\text{sr}} &= v \Delta t = (0.990c) \Delta t \\ &= (0.990)(299\,792\,458 \text{ m/s})(8.769 \times 10^{-7} \text{ s}) \\ &= 260 \text{ m.} \end{aligned} \quad (\text{Answer})$$

This is about seven times d_{cp} . Experiments like the one outlined here, which verify special relativity, became routine in physics laboratories decades ago. The engineering design and the construction of any scientific or medical facility that employs high-speed particles must take relativity into account.

37-6 The Relativity of Length

If you want to measure the length of a rod that is at rest with respect to you, you can—at your leisure—note the positions of its end points on a long stationary scale and subtract one reading from the other. If the rod is moving, however, you must note the positions of the end points *simultaneously* (in your reference frame) or your measurement cannot be called a length. Figure 37-7 suggests the difficulty of trying to measure the length of a moving penguin by locating its front and back at different times. Because simultaneity is relative and it enters into length measurements, length should also be a relative quantity. It is.

Fig. 37-7 If you want to measure the front-to-back length of a penguin while it is moving, you must mark the positions of its front and back simultaneously (in your reference frame), as in (a), rather than at different times, as in (b).



Let L_0 be the length of a rod that you measure when the rod is stationary (meaning you and it are in the same reference frame, the rod's rest frame). If, instead, there is relative motion at speed v between you and the rod *along the length of the rod*, then with simultaneous measurements you obtain a length L given by

$$L = L_0 \sqrt{1 - \beta^2} = \frac{L_0}{\gamma} \quad (\text{length contraction}). \quad (37-13)$$

Because the Lorentz factor γ is always greater than unity if there is relative motion, L is less than L_0 . The relative motion causes a *length contraction*, and L is called a *contracted length*. Because γ increases with speed v , the length contraction also increases with v .



The length L_0 of an object measured in the rest frame of the object is its **proper length** or **rest length**. Measurements of the length from any reference frame that is in relative motion *parallel* to that length are always less than the proper length.

Be careful: Length contraction occurs only along the direction of relative motion. Also, the length that is measured does not have to be that of an object like a rod or a circle. Instead, it can be the length (or distance) between two objects in the same rest frame—for example, the Sun and a nearby star (which are, at least approximately, at rest relative to each other).

Does a moving object *really shrink*? Reality is based on observations and measurements; if the results are always consistent and if no error can be determined, then what is observed and measured is real. In that sense, the object really does shrink. However, a more precise statement is that the object is *really measured* to shrink—motion affects that measurement and thus reality.

When you measure a contracted length for, say, a rod, what does an observer moving with the rod say of your measurement? To that observer, you did not locate the two ends of the rod simultaneously. (Recall that observers in motion relative to each other do not agree about simultaneity.) To the observer, you first located the rod's front end and then, slightly later, its rear end, and that is why you measured a length that is less than the proper length.

Proof of Eq. 37-13

Length contraction is a direct consequence of time dilation. Consider once more our two observers. This time, both Sally, seated on a train moving through a station, and Sam, again on the station platform, want to measure the length of the platform. Sam, using a tape measure, finds the length to be L_0 , a proper length because the platform is at rest with respect to him. Sam also notes that Sally, on the train, moves through this length in a time $\Delta t = L_0/v$, where v is the speed of the train; that is,

$$L_0 = v \Delta t \quad (\text{Sam}). \quad (37-14)$$

This time interval Δt is not a proper time interval because the two events that define it (Sally passes the back of the platform and Sally passes the front of the platform) occur at two different places, and therefore Sam must use two synchronized clocks to measure the time interval Δt .

For Sally, however, the platform is moving past her. She finds that the two events measured by Sam occur *at the same place* in her reference frame. She can time them with a single stationary clock, and so the interval Δt_0 that she measures is a proper time interval. To her, the length L of the platform is given by

$$L = v \Delta t_0 \quad (\text{Sally}). \quad (37-15)$$

If we divide Eq. 37-15 by Eq. 37-14 and apply Eq. 37-9, the time dilation equation, we have

$$\frac{L}{L_0} = \frac{v \Delta t_0}{v \Delta t} = \frac{1}{\gamma},$$

or

$$L = \frac{L_0}{\gamma}, \quad (37-16)$$

which is Eq. 37-13, the length contraction equation.

Sample Problem

Time dilation and length contraction as seen from each frame

In Fig. 37-8, Sally (at point A) and Sam's spaceship (of proper length $L_0 = 230$ m) pass each other with constant relative speed v . Sally measures a time interval of $3.57 \mu s$ for the ship to pass her (from the passage of point B in Fig. 37-8a to the passage of point C in Fig. 37-8b). In terms of c , what is the relative speed v between Sally and the ship?

KEY IDEAS

Let's assume that speed v is near c . Then:

1. This problem involves measurements made from two (inertial) reference frames, one attached to Sally and the other attached to Sam and his spaceship.
2. This problem also involves two events: the first is the passage of point B past Sally (Fig. 37-8a) and the second is the passage of point C past her (Fig. 37-8b).
3. From either reference frame, the other reference frame passes at speed v and moves a certain distance in the time interval between the two events:

$$v = \frac{\text{distance}}{\text{time interval}}. \quad (37-17)$$

Because speed v is assumed to be near the speed of light, we must be careful that the distance and the time interval in Eq. 37-17 are measured in the *same* reference frame. Otherwise, *speed* has no meaning.

Calculations: We are free to use either frame for the measurements. Because we know that the time interval Δt between the two events measured from Sally's frame is $3.57 \mu s$, let us also use the distance L between the two events measured from her frame. Equation 37-17 then becomes

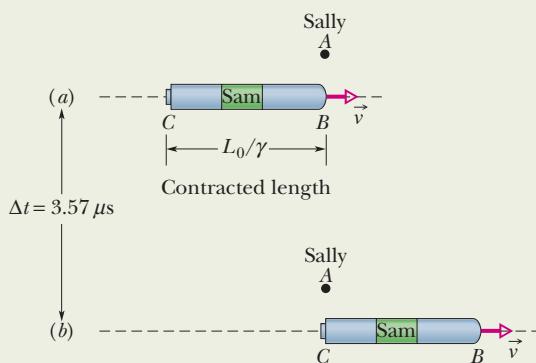
$$v = \frac{L}{\Delta t}. \quad (37-18)$$

We do not know L , but we can relate it to the given L_0 : The distance between the two events as measured from Sam's frame is the ship's proper length L_0 . Thus, the distance L measured from Sally's frame must be less than L_0 , as given by Eq. 37-13 ($L = L_0/\gamma$) for length contraction. Substituting L_0/γ for L in Eq. 37-18 and then substituting Eq. 37-8 for γ , we find

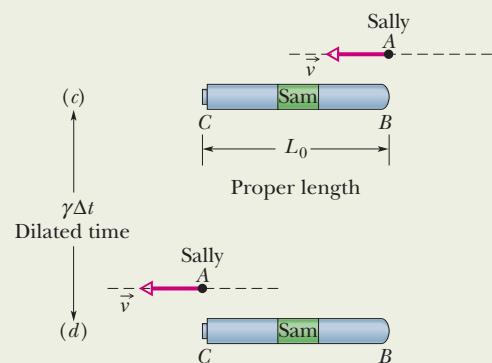
$$v = \frac{L_0/\gamma}{\Delta t} = \frac{L_0 \sqrt{1 - (v/c)^2}}{\Delta t}.$$

Fig. 37-8 (a)–(b)
Event 1 occurs when point B passes Sally (at point A) and event 2 occurs when point C passes her. (c)–(d) Event 1 occurs when Sally passes point B and event 2 occurs when she passes point C.

These are Sally's measurements, from her reference frame:



These are Sam's measurements, from his reference frame:



Solving this equation for v (notice that it is on the left and also buried in the Lorentz factor) leads us to

$$\begin{aligned} v &= \frac{L_0 c}{\sqrt{(c \Delta t)^2 + L_0^2}} \\ &= \frac{(230 \text{ m})c}{\sqrt{(299\,792\,458 \text{ m/s})^2(3.57 \times 10^{-6} \text{ s})^2 + (230 \text{ m})^2}} \\ &= 0.210c. \end{aligned} \quad (\text{Answer})$$

Thus, the relative speed between Sally and the ship is 21% of the speed of light.

Note that only the relative motion of Sally and Sam matters here; whether either is stationary relative to, say, a space station is irrelevant. In Figs. 37-8a and b we took Sally to be stationary, but we could instead have taken the ship to

be stationary, with Sally moving to the left past it. Event 1 is again when Sally and point *B* are aligned (Fig. 37-8c), and event 2 is again when Sally and point *C* are aligned (Fig. 37-8d). However, we are now using Sam's measurements. So the length between the two events in *his* frame is the proper length L_0 of the ship and the time interval between them is not Sally's measurement Δt but a dilated time interval $\gamma \Delta t$.

Substituting Sam's measurements into Eq. 37-17, we have

$$v = \frac{L_0}{\gamma \Delta t},$$

which is exactly what we found using Sally's measurements. Thus, we get the same result of $v = 0.210c$ with either set of measurements, *but we must be careful not to mix the measurements from the two frames*.

Sample Problem

Time dilation and length contraction in outrunning a supernova

Caught by surprise near a supernova, you race away from the explosion in your spaceship, hoping to outrun the high-speed material ejected toward you. Your Lorentz factor γ relative to the inertial reference frame of the local stars is 22.4.

(a) To reach a safe distance, you figure you need to cover $9.00 \times 10^{16} \text{ m}$ as measured in the reference frame of the local stars. How long will the flight take, as measured in that frame?

KEY IDEAS

From Chapter 2, for constant speed, we know that

$$\text{speed} = \frac{\text{distance}}{\text{time interval}}. \quad (37-19)$$

From Fig. 37-6, we see that because your Lorentz factor γ relative to the stars is 22.4 (large), your relative speed v is almost c —so close that we can approximate it as c . Then for speed $v \approx c$, we must be careful that the distance and the time interval in Eq. 37-19 are measured in the *same* reference frame.

Calculations: The given distance ($9.00 \times 10^{16} \text{ m}$) for the length of your travel path is measured in the reference frame of the stars, and the requested time interval Δt is to be measured in that same frame. Thus, we can write

$$\left(\frac{\text{time interval}}{\text{relative to stars}} \right) = \frac{\text{distance relative to stars}}{c}.$$

Then substituting the given distance, we find that

$$\begin{aligned} \left(\frac{\text{time interval}}{\text{relative to stars}} \right) &= \frac{9.00 \times 10^{16} \text{ m}}{299\,792\,458 \text{ m/s}} \\ &= 3.00 \times 10^8 \text{ s} = 9.51 \text{ y.} \end{aligned} \quad (\text{Answer})$$

(b) How long does that trip take according to you (in your reference frame)?

KEY IDEAS

- We now want the time interval measured in a different reference frame—namely, yours. Thus, we need to transform the data given in the reference frame of the stars to your frame.
- The given path length of $9.00 \times 10^{16} \text{ m}$, measured in the reference frame of the stars, is a proper length L_0 , because the two ends of the path are at rest in that frame. As observed from your reference frame, the stars' reference frame and those two ends of the path race past you at a relative speed of $v \approx c$.
- You measure a contracted length L_0/γ for the path, not the proper length L_0 .

Calculations: We can now rewrite Eq. 37-19 as

$$\left(\frac{\text{time interval}}{\text{relative to you}} \right) = \frac{\text{distance relative to you}}{c} = \frac{L_0/\gamma}{c}.$$

Substituting known data, we find

$$\begin{aligned} \left(\frac{\text{time interval}}{\text{relative to you}} \right) &= \frac{(9.00 \times 10^{16} \text{ m})/22.4}{299\,792\,458 \text{ m/s}} \\ &= 1.340 \times 10^7 \text{ s} = 0.425 \text{ y.} \end{aligned} \quad (\text{Answer})$$

In part (a) we found that the flight takes 9.51 y in the reference frame of the stars. However, here we find that it takes only 0.425 y in your frame, due to the relative motion and the resulting contracted length of the path.



Additional examples, video, and practice available at WileyPLUS

37-7 The Lorentz Transformation

Figure 37-9 shows inertial reference frame S' moving with speed v relative to frame S , in the common positive direction of their horizontal axes (marked x and x'). An observer in S reports spacetime coordinates x, y, z, t for an event, and an observer in S' reports x', y', z', t' for the same event. How are these sets of numbers related?

We claim at once (although it requires proof) that the y and z coordinates, which are perpendicular to the motion, are not affected by the motion; that is, $y = y'$ and $z = z'$. Our interest then reduces to the relation between x and x' and that between t and t' .

The Galilean Transformation Equations

Prior to Einstein's publication of his special theory of relativity, the four coordinates of interest were assumed to be related by the *Galilean transformation equations*:

$$\begin{aligned} x' &= x - vt && \text{(Galilean transformation equations;} \\ t' &= t && \text{approximately valid at low speeds).} \end{aligned} \quad (37-20)$$

(These equations are written with the assumption that $t = t' = 0$ when the origins of S and S' coincide.) You can verify the first equation with Fig. 37-9. The second equation effectively claims that time passes at the same rate for observers in both reference frames. That would have been so obviously true to a scientist prior to Einstein that it would not even have been mentioned. When speed v is small compared to c , Eqs. 37-20 generally work well.

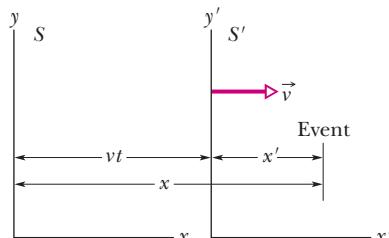


Fig. 37-9 Two inertial reference frames: frame S' has velocity \vec{v} relative to frame S .

The Lorentz Transformation Equations

We state without proof that the correct transformation equations, which remain valid for all speeds up to the speed of light, can be derived from the postulates of relativity. The results, called the **Lorentz transformation equations*** or sometimes (more loosely) just the Lorentz transformations, are

$$\begin{aligned} x' &= \gamma(x - vt), \\ y' &= y, \\ z' &= z, \\ t' &= \gamma(t - vx/c^2) \end{aligned} \quad \begin{array}{l} \text{(Lorentz transformation equations;} \\ \text{valid at all physically possible speeds).} \end{array} \quad (37-21)$$

(The equations are written with the assumption that $t = t' = 0$ when the origins of S and S' coincide.) Note that the spatial values x and the temporal values t are

* You may wonder why we do not call these the *Einstein transformation equations* (and why not the *Einstein factor* for γ). H. A. Lorentz actually derived these equations before Einstein did, but as the great Dutch physicist graciously conceded, he did not take the further bold step of interpreting these equations as describing the true nature of space and time. It is this interpretation, first made by Einstein, that is at the heart of relativity.

bound together in the first and last equations. This entanglement of space and time was a prime message of Einstein's theory, a message that was long rejected by many of his contemporaries.

It is a formal requirement of relativistic equations that they should reduce to familiar classical equations if we let c approach infinity. That is, if the speed of light were infinitely great, *all* finite speeds would be "low" and classical equations would never fail. If we let $c \rightarrow \infty$ in Eqs. 37-21, $\gamma \rightarrow 1$ and these equations reduce—as we expect—to the Galilean equations (Eqs. 37-20). You should check this.

Equations 37-21 are written in a form that is useful if we are given x and t and wish to find x' and t' . We may wish to go the other way, however. In that case we simply solve Eqs. 37-21 for x and t , obtaining

$$x = \gamma(x' + vt') \quad \text{and} \quad t = \gamma(t' + vx'/c^2). \quad (37-22)$$

Comparison shows that, starting from either Eqs. 37-21 or Eqs. 37-22, you can find the other set by interchanging primed and unprimed quantities and reversing the sign of the relative velocity v . (For example, if the S' frame has a positive velocity relative to an observer in the S frame as in Fig. 37-9, then the S frame has a *negative* velocity relative to an observer in the S' frame.)

Equations 37-21 and 37-22 relate the coordinates of a single event as seen by two observers. Sometimes we want to know not the coordinates of a single event but the differences between coordinates for a pair of events. That is, if we label our events 1 and 2, we may want to relate

$$\Delta x = x_2 - x_1 \quad \text{and} \quad \Delta t = t_2 - t_1,$$

as measured by an observer in S , and

$$\Delta x' = x'_2 - x'_1 \quad \text{and} \quad \Delta t' = t'_2 - t'_1,$$

as measured by an observer in S' .

Table 37-2 displays the Lorentz equations in difference form, suitable for analyzing pairs of events. The equations in the table were derived by simply substituting differences (such as Δx and $\Delta x'$) for the four variables in Eqs. 37-21 and 37-22.

Table 37-2

The Lorentz Transformation Equations for Pairs of Events

1. $\Delta x = \gamma(\Delta x' + v \Delta t')$	1'. $\Delta x' = \gamma(\Delta x - v \Delta t)$
2. $\Delta t = \gamma(\Delta t' + v \Delta x'/c^2)$	2'. $\Delta t' = \gamma(\Delta t - v \Delta x/c^2)$

$$\gamma = \frac{1}{\sqrt{1 - (v/c)^2}} = \frac{1}{\sqrt{1 - \beta^2}}$$

Frame S' moves at velocity v relative to frame S .

Be careful: When substituting values for these differences, you must be consistent and not mix the values for the first event with those for the second event. Also, if, say, Δx is a negative quantity, you must be certain to include the minus sign in a substitution.



CHECKPOINT 2

In Fig. 37-9, frame S' has velocity $0.90c$ relative to frame S . An observer in frame S' measures two events as occurring at the following spacetime coordinates: event Yellow at $(5.0 \text{ m}, 20 \text{ ns})$ and event Green at $(-2.0 \text{ m}, 45 \text{ ns})$. An observer in frame S wants to find the temporal separation $\Delta t_{GY} = t_G - t_Y$ between the events. (a) Which equation in Table 37-2 should be used? (b) Should $+0.90c$ or $-0.90c$ be substituted for v in the parentheses on the equation's right side and in the Lorentz factor γ ? What value should be substituted into the (c) first and (d) second term in the parentheses?

37-8 Some Consequences of the Lorentz Equations

Here we use the equations of Table 37-2 to affirm some of the conclusions that we reached earlier by arguments based directly on the postulates.

Simultaneity

Consider Eq. 2 of Table 37-2,

$$\Delta t = \gamma \left(\Delta t' + \frac{v \Delta x'}{c^2} \right). \quad (37-23)$$

If two events occur at different places in reference frame S' of Fig. 37-9, then $\Delta x'$ in this equation is not zero. It follows that even if the events are simultaneous in S' (thus $\Delta t' = 0$), they will not be simultaneous in frame S . (This is in accord with our conclusion in Section 37-4.) The time interval between the events in S will be

$$\Delta t = \gamma \frac{v \Delta x'}{c^2} \quad (\text{simultaneous events in } S').$$

Thus, the spatial separation $\Delta x'$ guarantees a temporal separation Δt .

Time Dilation

Suppose now that two events occur at the same place in S' (thus $\Delta x' = 0$) but at different times (thus $\Delta t' \neq 0$). Equation 37-23 then reduces to

$$\Delta t = \gamma \Delta t' \quad (\text{events in same place in } S'). \quad (37-24)$$

This confirms time dilation between frames S and S' . Moreover, because the two events occur at the same place in S' , the time interval $\Delta t'$ between them can be measured with a single clock, located at that place. Under these conditions, the measured interval is a proper time interval, and we can label it Δt_0 as we have previously labeled proper times. Thus, with that label Eq. 37-24 becomes

$$\Delta t = \gamma \Delta t_0 \quad (\text{time dilation}),$$

which is exactly Eq. 37-9, the time dilation equation. Thus, time dilation is a special case of the more general Lorentz equations.

Length Contraction

Consider Eq. 1' of Table 37-2,

$$\Delta x' = \gamma(\Delta x - v \Delta t). \quad (37-25)$$

If a rod lies parallel to the x and x' axes of Fig. 37-9 and is at rest in reference frame S' , an observer in S' can measure its length at leisure. One way to do so is by subtracting the coordinates of the end points of the rod. The value of $\Delta x'$ that is obtained will be the proper length L_0 of the rod because the measurements are made in a frame where the rod is at rest.

Suppose the rod is moving in frame S . This means that Δx can be identified as the length L of the rod in frame S only if the coordinates of the rod's end points are measured *simultaneously*—that is, if $\Delta t = 0$. If we put $\Delta x' = L_0$, $\Delta x = L$, and $\Delta t = 0$ in Eq. 37-25, we find

$$L = \frac{L_0}{\gamma} \quad (\text{length contraction}), \quad (37-26)$$

which is exactly Eq. 37-13, the length contraction equation. Thus, length contraction is a special case of the more general Lorentz equations.

Sample Problem

Lorentz transformations and reversing the sequence of events

An Earth starship has been sent to check an Earth outpost on the planet P1407, whose moon houses a battle group of the often hostile Reptilians. As the ship follows a straight-line course first past the planet and then past the moon, it detects a high-energy microwave burst at the Reptilian moon base and then, 1.10 s later, an explosion at the Earth outpost, which is 4.00×10^8 m from the Reptilian base as measured from the ship's reference frame. The Reptilians have obviously attacked the Earth outpost, and so the starship begins to prepare for a confrontation with them.

- (a) The speed of the ship relative to the planet and its moon is $0.980c$. What are the distance and time interval between the burst and the explosion as measured in the planet–moon frame (and thus according to the occupants of the stations)?

KEY IDEAS

- This problem involves measurements made from two reference frames, the planet–moon frame and the starship frame.
- We have two events: the burst and the explosion.
- We need to transform the given data as measured in the starship frame to the corresponding data as measured in the planet–moon frame.

Starship frame: Before we get to the transformation, we need to carefully choose our notation. We begin with a sketch of the situation as shown in Fig. 37-10. There, we have chosen the ship's frame S to be stationary and the planet–moon frame S' to be moving with positive velocity (rightward). (This is an arbitrary choice; we could, instead, have chosen the planet–moon frame to be stationary. Then we would redraw \vec{v} in Fig. 37-10 as being attached to the S frame and indicating leftward motion; v would then be a negative quantity. The results would be the same.) Let subscripts e and b represent the explosion and burst, respectively. Then the given data, all in the unprimed (starship) reference frame, are

$$\Delta x = x_e - x_b = +4.00 \times 10^8 \text{ m}$$

and $\Delta t = t_e - t_b = +1.10 \text{ s}$

Here, Δx is a positive quantity because in Fig. 37-10, the coordinate x_e for the explosion is greater than the coordinate x_b for the burst; Δt is also a positive quantity because the time t_e of the explosion is greater (later) than the time t_b of the burst.

Planet–moon frame: We seek $\Delta x'$ and $\Delta t'$, which we shall get by transforming the given S -frame data to the planet–moon frame S' . Because we are considering a pair of events, we choose transformation equations from Table 37-2—namely, Eqs. 1' and 2':

$$\Delta x' = \gamma(\Delta x - v \Delta t) \quad (37-27)$$

The relative motion alters the time intervals between events and maybe even their sequence.

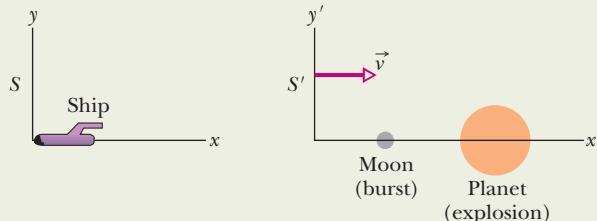


Fig. 37-10 A planet and its moon in reference frame S' move rightward with speed v relative to a starship in reference frame S .

and
$$\Delta t' = \gamma \left(\Delta t - \frac{v \Delta x}{c^2} \right). \quad (37-28)$$

Here, $v = +0.980c$ and the Lorentz factor is

$$\gamma = \frac{1}{\sqrt{1 - (v/c)^2}} = \frac{1}{\sqrt{1 - (+0.980c/c)^2}} = 5.0252.$$

Equation 37-27 then becomes

$$\begin{aligned} \Delta x' &= (5.0252)[4.00 \times 10^8 \text{ m} - (+0.980c)(1.10 \text{ s})] \\ &= 3.86 \times 10^8 \text{ m}, \end{aligned} \quad (\text{Answer})$$

and Eq. 37-28 becomes

$$\begin{aligned} \Delta t' &= (5.0252) \left[(1.10 \text{ s}) - \frac{(+0.980c)(4.00 \times 10^8 \text{ m})}{c^2} \right] \\ &= -1.04 \text{ s}. \end{aligned} \quad (\text{Answer})$$

(b) What is the meaning of the minus sign in the value for $\Delta t'$?

Reasoning: We must be consistent with the notation we set up in part (a). Recall how we originally defined the time interval between burst and explosion: $\Delta t = t_e - t_b = +1.10 \text{ s}$. To be consistent with that choice of notation, our definition of $\Delta t'$ must be $t'_e - t'_b$; thus, we have found that

$$\Delta t' = t'_e - t'_b = -1.04 \text{ s}.$$

The minus sign here tells us that $t'_b > t'_e$; that is, in the planet–moon reference frame, the burst occurred 1.04 s *after* the explosion, not 1.10 s *before* the explosion as detected in the ship frame.

(c) Did the burst cause the explosion, or vice versa?

KEY IDEA

The sequence of events measured in the planet–moon reference frame is the reverse of that measured in the ship frame. In either situation, if there is a causal relationship between the two events, information must travel from the location of one event to the location of the other to cause it.

Checking the speed: Let us check the required speed of the information. In the ship frame, this speed is

$$v_{\text{info}} = \frac{\Delta x}{\Delta t} = \frac{4.00 \times 10^8 \text{ m}}{1.10 \text{ s}} = 3.64 \times 10^8 \text{ m/s},$$

but that speed is impossible because it exceeds c . In the planet–moon frame, the speed comes out to be $3.70 \times 10^8 \text{ m/s}$, also impossible. Therefore, neither event could possibly have caused the other event; that is, they are *unrelated* events. Thus, the starship should not confront the Reptilians.



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37-9 The Relativity of Velocities

Here we wish to use the Lorentz transformation equations to compare the velocities that two observers in different inertial reference frames S and S' would measure for the same moving particle. Let S' move with velocity v relative to S .

Suppose that the particle, moving with constant velocity parallel to the x and x' axes in Fig. 37-11, sends out two signals as it moves. Each observer measures the space interval and the time interval between these two events. These four measurements are related by Eqs. 1 and 2 of Table 37-2,

$$\Delta x = \gamma(\Delta x' + v \Delta t')$$

and

$$\Delta t = \gamma\left(\Delta t' + \frac{v \Delta x'}{c^2}\right).$$

If we divide the first of these equations by the second, we find

$$\frac{\Delta x}{\Delta t} = \frac{\Delta x' + v \Delta t'}{\Delta t' + v \Delta x'/c^2}.$$

Dividing the numerator and denominator of the right side by $\Delta t'$, we find

$$\frac{\Delta x}{\Delta t} = \frac{\Delta x'/\Delta t' + v}{1 + v(\Delta x'/\Delta t')/c^2}.$$

However, in the differential limit, $\Delta x/\Delta t$ is u , the velocity of the particle as measured in S , and $\Delta x'/\Delta t'$ is u' , the velocity of the particle as measured in S' . Then we have, finally,

$$u = \frac{u' + v}{1 + u'v/c^2} \quad (\text{relativistic velocity transformation}) \quad (37-29)$$

as the relativistic velocity transformation equation. This equation reduces to the classical, or Galilean, velocity transformation,

$$u = u' + v \quad (\text{classical velocity transformation}), \quad (37-30)$$

when we apply the formal test of letting $c \rightarrow \infty$. In other words, Eq. 37-29 is correct for all physically possible speeds, but Eq. 37-30 is approximately correct for speeds much less than c .

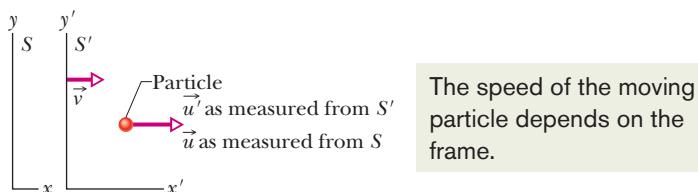


Fig. 37-11 Reference frame S' moves with velocity \vec{v} relative to frame S . A particle has velocity \vec{u}' relative to reference frame S' and velocity \vec{u} relative to reference frame S .

37-10 Doppler Effect for Light

In Section 17-9 we discussed the Doppler effect (a shift in detected frequency) for sound waves traveling in air. For such waves, the Doppler effect depends on two velocities—namely, the velocities of the source and detector with respect to the air. Air is the medium that transmits the waves.

That is not the situation with light waves, for they (and other electromagnetic waves) require no medium, being able to travel even through vacuum. The Doppler effect for light waves depends on only one velocity, the relative velocity \vec{v} between source and detector, as measured from the reference frame of either. Let f_0 represent the **proper frequency** of the source—that is, the frequency that is measured by an observer in the rest frame of the source. Let f represent the frequency detected by an observer moving with velocity \vec{v} relative to that rest frame. Then, when the direction of \vec{v} is directly away from the source,

$$f = f_0 \sqrt{\frac{1 - \beta}{1 + \beta}} \quad (\text{source and detector separating}), \quad (37-31)$$

where $\beta = v/c$. When the direction of \vec{v} is directly toward the source, we must change the signs in front of both β symbols in Eq. 37-31.

Low-Speed Doppler Effect

For low speeds ($\beta \ll 1$), Eq. 37-31 can be expanded in a power series in β and approximated as

$$f = f_0(1 - \beta + \frac{1}{2}\beta^2) \quad (\text{source and detector separating}, \beta \ll 1). \quad (37-32)$$

The corresponding low-speed equation for the Doppler effect with sound waves (or any waves except light waves) has the same first two terms but a different coefficient in the third term. Thus, the relativistic effect for low-speed light sources and detectors shows up only with the β^2 term.

A police radar unit employs the Doppler effect with microwaves to measure the speed v of a car. A source in the radar unit emits a microwave beam at a certain (proper) frequency f_0 along the road. A car that is moving toward the unit intercepts that beam but at a frequency that is shifted upward by the Doppler effect due to the car's motion toward the radar unit. The car reflects the beam back toward the radar unit. Because the car is moving toward the radar unit, the detector in the unit intercepts a reflected beam that is further shifted up in frequency. The unit compares that detected frequency with f_0 and computes the speed v of the car.

Astronomical Doppler Effect

In astronomical observations of stars, galaxies, and other sources of light, we can determine how fast the sources are moving, either directly away from us or directly toward us, by measuring the *Doppler shift* of the light that reaches us. If a certain star were at rest relative to us, we would detect light from it with a certain proper frequency f_0 . However, if the star is moving either directly away from us or directly toward us, the light we detect has a frequency f that is shifted from f_0 by the Doppler effect. This Doppler shift is due only to the *radial* motion of the star (its motion directly toward us or away from us), and the speed we can determine by measuring this Doppler shift is only the *radial speed* v of the star—that is, only the radial component of the star's velocity relative to us.

Suppose a star (or any other light source) moves away from us with a radial speed v that is low enough (β is small enough) for us to neglect the β^2 term in Eq. 37-32. Then we have

$$f = f_0(1 - \beta). \quad (37-33)$$

37-10 DOPPLER EFFECT FOR LIGHT

1041

Because astronomical measurements involving light are usually done in wavelengths rather than frequencies, let us replace f with c/λ and f_0 with c/λ_0 , where λ is the measured wavelength and λ_0 is the **proper wavelength** (the wavelength associated with f_0). We then have

$$\frac{c}{\lambda} = \frac{c}{\lambda_0} (1 - \beta),$$

or $\lambda = \lambda_0(1 - \beta)^{-1}$. (37-34)

Because we assume β is small, we can expand $(1 - \beta)^{-1}$ in a power series. Doing so and retaining only the first power of β , we have

$$\lambda = \lambda_0(1 + \beta),$$

or $\beta = \frac{\lambda - \lambda_0}{\lambda_0}$. (37-35)

Replacing β with v/c and $\lambda - \lambda_0$ with $|\Delta\lambda|$ leads to

$$v = \frac{|\Delta\lambda|}{\lambda_0} c \quad (\text{radial speed of light source, } v \ll c). \quad (37-36)$$

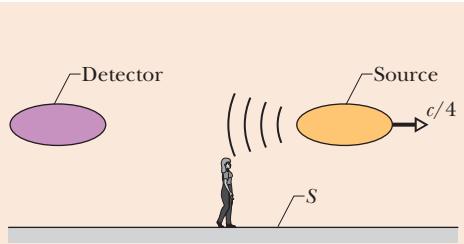
The difference $\Delta\lambda$ is the *wavelength Doppler shift* of the light source. We enclose it with an absolute sign so that we always have a magnitude of the shift.

Equation 37-36 is an approximation that can be applied only when $v \ll c$. Under that condition, Eq. 37-36 can be applied whether the light source is moving toward or away from us. If it is moving away from us, then λ is longer than λ_0 , $\Delta\lambda$ is positive, and the Doppler shift is called a *red shift*. (The term *red* does not mean the detected light is red or even visible. It merely serves as a memory device because red is at the *long* wavelength end of the visible spectrum. Thus λ is longer than λ_0 .) If the light source is moving toward us, then λ is shorter than λ_0 , $\Delta\lambda$ is negative, and the Doppler shift is called a *blue shift*.



CHECKPOINT 3

The figure shows a source that emits light of proper frequency f_0 while moving directly toward the right with speed $c/4$ as measured from reference frame S . The figure also shows a light detector, which measures a frequency $f > f_0$ for the emitted light. (a) Is the detector moving toward the left or the right? (b) Is the speed of the detector as measured from reference frame S more than $c/4$, less than $c/4$, or equal to $c/4$?



Transverse Doppler Effect

So far, we have discussed the Doppler effect, here and in Chapter 17, only for situations in which the source and the detector move either directly toward or directly away from each other. Figure 37-12 shows a different arrangement, in which a source S moves past a detector D . When S reaches point P , the velocity of S is perpendicular to the line joining P and D , and at that instant S is moving neither toward nor away from D . If the source is emitting sound waves of frequency f_0 , D detects that frequency (with no Doppler effect) when it intercepts the waves that were emitted at point P . However, if the source is emitting light waves, there is still a Doppler effect, called the **transverse Doppler effect**. In this situation, the detected frequency of the light emitted when the source is at point P is

$$f = f_0 \sqrt{1 - \beta^2} \quad (\text{transverse Doppler effect}). \quad (37-37)$$

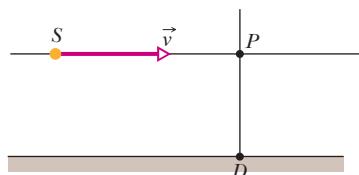


Fig. 37-12 A light source S travels with velocity \vec{v} past a detector at D . The special theory of relativity predicts a transverse Doppler effect as the source passes through point P , where the direction of travel is perpendicular to the line extending through D . Classical theory predicts no such effect.

For low speeds ($\beta \ll 1$), Eq. 37-37 can be expanded in a power series in β and approximated as

$$f = f_0(1 - \frac{1}{2}\beta^2) \quad (\text{low speeds}). \quad (37-38)$$

Here the first term is what we would expect for sound waves, and again the relativistic effect for low-speed light sources and detectors appears with the β^2 term.

In principle, a police radar unit can determine the speed of a car even when the path of the radar beam is perpendicular (transverse) to the path of the car. However, Eq. 37-38 tells us that because β is small even for a fast car, the relativistic term $\beta^2/2$ in the transverse Doppler effect is extremely small. Thus, $f \approx f_0$ and the radar unit computes a speed of zero.

The transverse Doppler effect is really another test of time dilation. If we rewrite Eq. 37-37 in terms of the period T of oscillation of the emitted light wave instead of the frequency, we have, because $T = 1/f$,

$$T = \frac{T_0}{\sqrt{1 - \beta^2}} = \gamma T_0, \quad (37-39)$$

in which T_0 ($= 1/f_0$) is the **proper period** of the source. As comparison with Eq. 37-9 shows, Eq. 37-39 is simply the time dilation formula because a period is a time interval.

37-11 A New Look at Momentum

Suppose that a number of observers, each in a different inertial reference frame, watch an isolated collision between two particles. In classical mechanics, we have seen that—even though the observers measure different velocities for the colliding particles—they all find that the law of conservation of momentum holds. That is, they find that the total momentum of the system of particles after the collision is the same as it was before the collision.

How is this situation affected by relativity? We find that if we continue to define the momentum \vec{p} of a particle as $m\vec{v}$, the product of its mass and its velocity, total momentum is *not* conserved for the observers in different inertial frames. We have two choices: (1) Give up the law of conservation of momentum or (2) see whether we can refine our definition of momentum in some new way so that the law of conservation of momentum still holds. The correct choice is the second one.

Consider a particle moving with constant speed v in the positive direction of an x axis. Classically, its momentum has magnitude

$$p = mv = m \frac{\Delta x}{\Delta t} \quad (\text{classical momentum}), \quad (37-40)$$

in which Δx is the distance it travels in time Δt . To find a relativistic expression for momentum, we start with the new definition

$$p = m \frac{\Delta x}{\Delta t_0}.$$

Here, as before, Δx is the distance traveled by a moving particle as viewed by an observer watching that particle. However, Δt_0 is the time required to travel that distance, measured not by the observer watching the moving particle but by an observer moving with the particle. The particle is at rest with respect to this second observer; thus that measured time is a proper time.

Using the time dilation formula, $\Delta t = \gamma \Delta t_0$ (Eq. 37-9), we can then write

$$p = m \frac{\Delta x}{\Delta t_0} = m \frac{\Delta x}{\Delta t} \frac{\Delta t}{\Delta t_0} = m \frac{\Delta x}{\Delta t} \gamma.$$

However, since $\Delta x/\Delta t$ is just the particle velocity v , we have

$$p = \gamma mv \quad (\text{momentum}). \quad (37-41)$$

Note that this differs from the classical definition of Eq. 37-40 only by the Lorentz factor γ . However, that difference is important: Unlike classical momentum, relativistic momentum approaches an infinite value as v approaches c .

We can generalize the definition of Eq. 37-41 to vector form as

$$\vec{p} = \gamma m \vec{v} \quad (\text{momentum}). \quad (37-42)$$

This equation gives the correct definition of momentum for all physically possible speeds. For a speed much less than c , it reduces to the classical definition of momentum ($\vec{p} = m\vec{v}$).

37-12 A New Look at Energy

Mass Energy

The science of chemistry was initially developed with the assumption that in chemical reactions, energy and mass are conserved separately. In 1905, Einstein showed that as a consequence of his theory of special relativity, mass can be considered to be another form of energy. Thus, the law of conservation of energy is really the law of conservation of mass–energy.

In a *chemical reaction* (a process in which atoms or molecules interact), the amount of mass that is transferred into other forms of energy (or vice versa) is such a tiny fraction of the total mass involved that there is no hope of measuring the mass change with even the best laboratory balances. Mass and energy truly *seem* to be separately conserved. However, in a *nuclear reaction* (in which nuclei or fundamental particles interact), the energy released is often about a million times greater than in a chemical reaction, and the change in mass can easily be measured.

An object's mass m and the equivalent energy E_0 are related by

$$E_0 = mc^2, \quad (37-43)$$

which, without the subscript 0, is the best-known science equation of all time. This energy that is associated with the mass of an object is called **mass energy** or **rest energy**. The second name suggests that E_0 is an energy that the object has even when it is at rest, simply because it has mass. (If you continue your study of physics beyond this book, you will see more refined discussions of the relation between mass and energy. You might even encounter disagreements about just what that relation is and means.)

Table 37-3 shows the (approximate) mass energy, or rest energy, of a few objects. The mass energy of, say, a U.S. penny is enormous; the equivalent amount

Table 37-3

The Energy Equivalents of a Few Objects

Object	Mass (kg)	Energy Equivalent	
Electron	$\approx 9.11 \times 10^{-31}$	$\approx 8.19 \times 10^{-14}$ J	(≈ 511 keV)
Proton	$\approx 1.67 \times 10^{-27}$	$\approx 1.50 \times 10^{-10}$ J	(≈ 938 MeV)
Uranium atom	$\approx 3.95 \times 10^{-25}$	$\approx 3.55 \times 10^{-8}$ J	(≈ 225 GeV)
Dust particle	$\approx 1 \times 10^{-13}$	$\approx 1 \times 10^4$ J	(≈ 2 kcal)
U.S. penny	$\approx 3.1 \times 10^{-3}$	$\approx 2.8 \times 10^{14}$ J	(≈ 78 GW·h)

of electrical energy would cost well over a million dollars. On the other hand, the entire annual U.S. electrical energy production corresponds to a mass of only a few hundred kilograms of matter (stones, burritos, or anything else).

In practice, SI units are rarely used with Eq. 37-43 because they are too large to be convenient. Masses are usually measured in atomic mass units, where

$$1 \text{ u} = 1.660\,538\,86 \times 10^{-27} \text{ kg}, \quad (37-44)$$

and energies are usually measured in electron-volts or multiples of it, where

$$1 \text{ eV} = 1.602\,176\,462 \times 10^{-19} \text{ J}. \quad (37-45)$$

In the units of Eqs. 37-44 and 37-45, the multiplying constant c^2 has the values

$$\begin{aligned} c^2 &= 9.314\,940\,13 \times 10^8 \text{ eV/u} = 9.314\,940\,13 \times 10^5 \text{ keV/u} \\ &= 931.494\,013 \text{ MeV/u}. \end{aligned} \quad (37-46)$$

Total Energy

Equation 37-43 gives, for any object, the mass energy E_0 that is associated with the object's mass m , regardless of whether the object is at rest or moving. If the object is moving, it has additional energy in the form of kinetic energy K . If we assume that the object's potential energy is zero, then its total energy E is the sum of its mass energy and its kinetic energy:

$$E = E_0 + K = mc^2 + K. \quad (37-47)$$

Although we shall not prove it, the total energy E can also be written as

$$E = \gamma mc^2, \quad (37-48)$$

where γ is the Lorentz factor for the object's motion.

Since Chapter 7, we have discussed many examples involving changes in the total energy of a particle or a system of particles. However, we did not include mass energy in the discussions because the changes in mass energy were either zero or small enough to be neglected. The law of conservation of total energy still applies when changes in mass energy are significant. Thus, regardless of what happens to the mass energy, the following statement from Section 8-8 is still true:



The total energy E of an *isolated system* cannot change.

For example, if the total mass energy of two interacting particles in an isolated system decreases, some other type of energy in the system must increase because the total energy cannot change.

In a system undergoing a chemical or nuclear reaction, a change in the total mass energy of the system due to the reaction is often given as a Q value. The Q value for a reaction is obtained from the relation

$$\left(\begin{array}{l} \text{system's initial} \\ \text{total mass energy} \end{array} \right) = \left(\begin{array}{l} \text{system's final} \\ \text{total mass energy} \end{array} \right) + Q$$

or $E_{0i} = E_{0f} + Q. \quad (37-49)$

Using Eq. 37-43 ($E_0 = mc^2$), we can rewrite this in terms of the initial *total* mass M_i and the final *total* mass M_f as

$$M_i c^2 = M_f c^2 + Q$$

or $Q = M_i c^2 - M_f c^2 = -\Delta M c^2, \quad (37-50)$

where the change in mass due to the reaction is $\Delta M = M_f - M_i$.

If a reaction results in the transfer of energy from mass energy to, say, kinetic energy of the reaction products, the system's total mass energy E_0 (and total mass M) decreases and Q is positive. If, instead, a reaction requires that energy be transferred to mass energy, the system's total mass energy E_0 (and its total mass M) increases and Q is negative.

For example, suppose two hydrogen nuclei undergo a *fusion reaction* in which they join together to form a single nucleus and release two particles in the process. The total mass energy (and total mass) of the resultant single nucleus and two released particles is less than the total mass energy (and total mass) of the initial hydrogen nuclei. Thus, the Q of the fusion reaction is positive, and energy is said to be *released* (transferred from mass energy) by the reaction. This release is important to you because the fusion of hydrogen nuclei in the Sun is one part of the process that results in sunshine on Earth and makes life here possible.

Kinetic Energy

In Chapter 7 we defined the kinetic energy K of an object of mass m moving at speed v well below c to be

$$K = \frac{1}{2}mv^2. \quad (37-51)$$

However, this classical equation is only an approximation that is good enough when the speed is well below the speed of light.

Let us now find an expression for kinetic energy that is correct for *all* physically possible speeds, including speeds close to c . Solving Eq. 37-47 for K and then substituting for E from Eq. 37-48 lead to

$$\begin{aligned} K &= E - mc^2 = \gamma mc^2 - mc^2 \\ &= mc^2(\gamma - 1) \quad (\text{kinetic energy}), \end{aligned} \quad (37-52)$$

where $\gamma (= 1/\sqrt{1 - (v/c)^2})$ is the Lorentz factor for the object's motion.

Figure 37-13 shows plots of the kinetic energy of an electron as calculated with the correct definition (Eq. 37-52) and the classical approximation (Eq. 37-51), both as functions of v/c . Note that on the left side of the graph the two plots coincide; this is the part of the graph—at lower speeds—where we have calculated kinetic energies so far in this book. This part of the graph tells us that we have been justified in calculating kinetic energy with the classical expression of Eq. 37-51. However, on the right side of the graph—at speeds near c —the two plots differ significantly. As v/c approaches 1.0, the plot for the classical definition of kinetic energy increases only moderately while the plot for the correct definition of kinetic energy increases dramatically, approaching an infinite value as v/c approaches 1.0. Thus, when an object's speed v is near c , we *must* use Eq. 37-52 to calculate its kinetic energy.

Figure 37-13 also tells us something about the work we must do on an object to increase its speed by, say, 1%. The required work W is equal to the resulting change ΔK in the object's kinetic energy. If the change is to occur on the low-speed, left side of Fig. 37-13, the required work might be modest. However, if the change is to occur on the high-speed, right side of Fig. 37-13, the required work could be enormous because the kinetic energy K increases so rapidly there with an increase in speed v . To increase an object's speed to c would require, in principle, an infinite amount of energy; thus, doing so is impossible.

The kinetic energies of electrons, protons, and other particles are often stated with the unit electron-volt or one of its multiples used as an adjective. For example, an electron with a kinetic energy of 20 MeV may be described as a 20 MeV electron.

As v/c approaches 1.0, the actual kinetic energy approaches infinity.

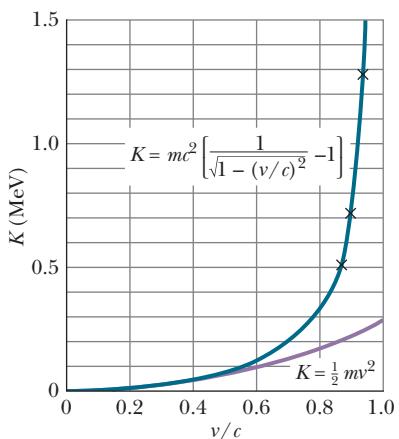


Fig. 37-13 The relativistic (Eq. 37-52) and classical (Eq. 37-51) equations for the kinetic energy of an electron, plotted as a function of v/c , where v is the speed of the electron and c is the speed of light. Note that the two curves blend together at low speeds and diverge widely at high speeds. Experimental data (at the \times marks) show that at high speeds the relativistic curve agrees with experiment but the classical curve does not.

This might help you to remember the relations.

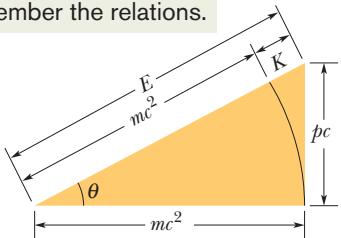


Fig. 37-14 A useful memory diagram for the relativistic relations among the total energy E , the rest energy or mass energy mc^2 , the kinetic energy K , and the momentum magnitude p .

Momentum and Kinetic Energy

In classical mechanics, the momentum p of a particle is mv and its kinetic energy K is $\frac{1}{2}mv^2$. If we eliminate v between these two expressions, we find a direct relation between momentum and kinetic energy:

$$p^2 = 2Km \quad (\text{classical}). \quad (37-53)$$

We can find a similar connection in relativity by eliminating v between the relativistic definition of momentum (Eq. 37-41) and the relativistic definition of kinetic energy (Eq. 37-52). Doing so leads, after some algebra, to

$$(pc)^2 = K^2 + 2Kmc^2. \quad (37-54)$$

With the aid of Eq. 37-47, we can transform Eq. 37-54 into a relation between the momentum p and the total energy E of a particle:

$$E^2 = (pc)^2 + (mc^2)^2. \quad (37-55)$$

The right triangle of Fig. 37-14 can help you keep these useful relations in mind. You can also show that, in that triangle,

$$\sin \theta = \beta \quad \text{and} \quad \cos \theta = 1/\gamma. \quad (37-56)$$

With Eq. 37-55 we can see that the product pc must have the same unit as energy E ; thus, we can express the unit of momentum p as an energy unit divided by c , usually as MeV/c or GeV/c in fundamental particle physics.



CHECKPOINT 4

Are (a) the kinetic energy and (b) the total energy of a 1 GeV electron more than, less than, or equal to those of a 1 GeV proton?

Sample Problem

Energy and momentum of a relativistic electron

- (a) What is the total energy E of a 2.53 MeV electron?

KEY IDEA

From Eq. 37-47, the total energy E is the sum of the electron's mass energy (or rest energy) mc^2 and its kinetic energy:

$$E = mc^2 + K. \quad (37-57)$$

Calculations: The adjective "2.53 MeV" in the problem statement means that the electron's kinetic energy is 2.53 MeV. To evaluate the electron's mass energy mc^2 , we substitute the electron's mass m from Appendix B, obtaining

$$\begin{aligned} mc^2 &= (9.109 \times 10^{-31} \text{ kg})(299\,792\,458 \text{ m/s})^2 \\ &= 8.187 \times 10^{-14} \text{ J}. \end{aligned}$$

Then dividing this result by $1.602 \times 10^{-13} \text{ J/MeV}$ gives us 0.511 MeV as the electron's mass energy (confirming the value in Table 37-3). Equation 37-57 then yields

$$E = 0.511 \text{ MeV} + 2.53 \text{ MeV} = 3.04 \text{ MeV}. \quad (\text{Answer})$$

- (b) What is the magnitude p of the electron's momentum, in the unit MeV/c? (Note that c is the symbol for the speed of light and not itself a unit.)

KEY IDEA

We can find p from the total energy E and the mass energy mc^2 via Eq. 37-55,

$$E^2 = (pc)^2 + (mc^2)^2.$$

Calculations: Solving for pc gives us

$$\begin{aligned} pc &= \sqrt{E^2 - (mc^2)^2} \\ &= \sqrt{(3.04 \text{ MeV})^2 - (0.511 \text{ MeV})^2} = 3.00 \text{ MeV}. \end{aligned}$$

Finally, dividing both sides by c we find

$$p = 3.00 \text{ MeV}/c. \quad (\text{Answer})$$



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Sample Problem**Energy and an astounding discrepancy in travel time**

The most energetic proton ever detected in the cosmic rays coming to Earth from space had an astounding kinetic energy of 3.0×10^{20} eV (enough energy to warm a teaspoon of water by a few degrees).

- (a) What were the proton's Lorentz factor γ and speed v (both relative to the ground-based detector)?

KEY IDEAS

(1) The proton's Lorentz factor γ relates its total energy E to its mass energy mc^2 via Eq. 37-48 ($E = \gamma mc^2$). (2) The proton's total energy is the sum of its mass energy mc^2 and its (given) kinetic energy K .

Calculations: Putting these ideas together we have

$$\gamma = \frac{E}{mc^2} = \frac{mc^2 + K}{mc^2} = 1 + \frac{K}{mc^2}. \quad (37-58)$$

From Table 37-3, the proton's mass energy mc^2 is 938 MeV. Substituting this and the given kinetic energy into Eq. 37-58, we obtain

$$\begin{aligned}\gamma &= 1 + \frac{3.0 \times 10^{20} \text{ eV}}{938 \times 10^6 \text{ eV}} \\ &= 3.198 \times 10^{11} \approx 3.2 \times 10^{11}. \quad (\text{Answer})\end{aligned}$$

This computed value for γ is so large that we cannot use the definition of γ (Eq. 37-8) to find v . Try it; your calculator will tell you that β is effectively equal to 1 and thus that v is effectively equal to c . Actually, v is almost c , but we want a more accurate answer, which we can obtain by first solving Eq. 37-8 for $1 - \beta$. To begin we write

$$\gamma = \frac{1}{\sqrt{1 - \beta^2}} = \frac{1}{\sqrt{(1 - \beta)(1 + \beta)}} \approx \frac{1}{\sqrt{2(1 - \beta)}},$$

where we have used the fact that β is so close to unity that $1 + \beta$ is very close to 2. (We can round off the sum of two very close numbers but not their difference.) The velocity we seek is contained in the $1 - \beta$ term. Solving for $1 - \beta$ then yields

$$\begin{aligned}1 - \beta &= \frac{1}{2\gamma^2} = \frac{1}{(2)(3.198 \times 10^{11})^2} \\ &= 4.9 \times 10^{-24} \approx 5 \times 10^{-24}.\end{aligned}$$

Thus, $\beta = 1 - 5 \times 10^{-24}$
and, since $v = \beta c$,

$$v \approx 0.999\ 999\ 999\ 999\ 999\ 999\ 999\ 995c. \quad (\text{Answer})$$

- (b) Suppose that the proton travels along a diameter of the Milky Way galaxy (9.8×10^4 ly). Approximately how long does the proton take to travel that diameter as measured from the common reference frame of Earth and the Galaxy?

Reasoning: We just saw that this *ultrarelativistic* proton is traveling at a speed barely less than c . By the definition of light-year, light takes 1 y to travel a distance of 1 ly, and so light should take 9.8×10^4 y to travel 9.8×10^4 ly, and this proton should take almost the same time. Thus, from our Earth–Milky Way reference frame, the proton's trip takes

$$\Delta t = 9.8 \times 10^4 \text{ y.} \quad (\text{Answer})$$

- (c) How long does the trip take as measured in the reference frame of the proton?

KEY IDEAS

1. This problem involves measurements made from two (inertial) reference frames: one is the Earth–Milky Way frame and the other is attached to the proton.
2. This problem also involves two events: the first is when the proton passes one end of the diameter along the Galaxy, and the second is when it passes the opposite end.
3. The time interval between those two events as measured in the proton's reference frame is the proper time interval Δt_0 because the events occur at the same location in that frame—namely, at the proton itself.
4. We can find the proper time interval Δt_0 from the time interval Δt measured in the Earth–Milky Way frame by using Eq. 37-9 ($\Delta t = \gamma \Delta t_0$) for time dilation. (Note that we can use that equation because one of the time measures is a proper time. However, we get the same relation if we use a Lorentz transformation.)

Calculation: Solving Eq. 37-9 for Δt_0 and substituting γ from (a) and Δt from (b), we find

$$\begin{aligned}\Delta t_0 &= \frac{\Delta t}{\gamma} = \frac{9.8 \times 10^4 \text{ y}}{3.198 \times 10^{11}} \\ &= 3.06 \times 10^{-7} \text{ y} = 9.7 \text{ s.} \quad (\text{Answer})\end{aligned}$$

In our frame, the trip takes 98 000 y. In the proton's frame, it takes 9.7 s! As promised at the start of this chapter, relative motion can alter the rate at which time passes, and we have here an extreme example.



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REVIEW & SUMMARY

The Postulates Einstein's **special theory of relativity** is based on two postulates:

1. The laws of physics are the same for observers in all inertial reference frames. No one frame is preferred over any other.
2. The speed of light in vacuum has the same value c in all directions and in all inertial reference frames.

The speed of light c in vacuum is an ultimate speed that cannot be exceeded by any entity carrying energy or information.

Coordinates of an Event Three space coordinates and one time coordinate specify an **event**. One task of special relativity is to relate these coordinates as assigned by two observers who are in uniform motion with respect to each other.

Simultaneous Events If two observers are in relative motion, they will not, in general, agree as to whether two events are simultaneous. If one of the observers finds two events at different locations to be simultaneous, the other will not, and conversely. Simultaneity is *not* an absolute concept but a relative one, depending on the motion of the observer. The relativity of simultaneity is a direct consequence of the finite ultimate speed c .

Time Dilation If two successive events occur at the same place in an inertial reference frame, the time interval Δt_0 between them, measured on a single clock where they occur, is the **proper time** between the events. *Observers in frames moving relative to that frame will measure a larger value for this interval.* For an observer moving with relative speed v , the measured time interval is

$$\begin{aligned}\Delta t &= \frac{\Delta t_0}{\sqrt{1 - (v/c)^2}} = \frac{\Delta t_0}{\sqrt{1 - \beta^2}} \\ &= \gamma \Delta t_0 \quad (\text{time dilation}).\end{aligned}\quad (37-7 \text{ to } 37-9)$$

Here $\beta = v/c$ is the **speed parameter** and $\gamma = 1/\sqrt{1 - \beta^2}$ is the **Lorentz factor**. An important result of time dilation is that moving clocks run slow as measured by an observer at rest.

Length Contraction The length L_0 of an object measured by an observer in an inertial reference frame in which the object is at rest is called its **proper length**. *Observers in frames moving relative to that frame and parallel to that length will measure a shorter length.* For an observer moving with relative speed v , the measured length is

$$L = L_0 \sqrt{1 - \beta^2} = \frac{L_0}{\gamma} \quad (\text{length contraction}). \quad (37-13)$$

The Lorentz Transformation The *Lorentz transformation* equations relate the spacetime coordinates of a single event as seen by observers in two inertial frames, S and S' , where S' is moving relative to S with velocity v in the positive x and x' direction. The four coordinates are related by

$$\begin{aligned}x' &= \gamma(x - vt), \\ y' &= y, \\ z' &= z, \\ t' &= \gamma(t - vx/c^2)\end{aligned}\quad (37-21)$$

Relativity of Velocities When a particle is moving with speed u' in the positive x' direction in an inertial reference frame S' that itself is moving with speed v parallel to the x direction of a second inertial frame S , the speed u of the particle as measured in S is

$$u = \frac{u' + v}{1 + u'v/c^2} \quad (\text{relativistic velocity}). \quad (37-29)$$

Relativistic Doppler Effect If a source emitting light waves of frequency f_0 moves directly away from a detector with relative radial speed v (and speed parameter $\beta = v/c$), the frequency f measured by the detector is

$$f = f_0 \sqrt{\frac{1 - \beta}{1 + \beta}}. \quad (37-31)$$

The motion results in a decrease in the detected frequency and (as more commonly measured) an increase in the wavelength (red shift). If the source moves directly toward the detector, the signs in Eq. 37-31 are reversed. Now the motion results in an increase in frequency and a decrease in wavelength (blue shift).

For astronomical observations, the Doppler effect is measured in wavelengths. For speeds much less than c , Eq. 37-31 leads to

$$v = \frac{|\Delta\lambda|}{\lambda_0} c, \quad (37-36)$$

where $\Delta\lambda (= \lambda - \lambda_0)$ is the *Doppler shift* in wavelength due to the motion.

Transverse Doppler Effect If the relative motion of the light source is perpendicular to a line joining the source and detector, the Doppler frequency formula is

$$f = f_0 \sqrt{1 - \beta^2}. \quad (37-37)$$

This **transverse Doppler effect** is due to time dilation.

Momentum and Energy The following definitions of linear momentum \vec{p} , kinetic energy K , and total energy E for a particle of mass m are valid at any physically possible speed:

$$\vec{p} = \gamma m \vec{v} \quad (\text{momentum}), \quad (37-42)$$

$$E = mc^2 + K = \gamma mc^2 \quad (\text{total energy}), \quad (37-47, 37-48)$$

$$K = mc^2(\gamma - 1) \quad (\text{kinetic energy}). \quad (37-52)$$

Here γ is the Lorentz factor for the particle's motion, and mc^2 is the *mass energy*, or *rest energy*, associated with the mass of the particle. These equations lead to the relationships

$$(pc)^2 = K^2 + 2Kmc^2, \quad (37-54)$$

and $E^2 = (pc)^2 + (mc^2)^2. \quad (37-55)$

When a system of particles undergoes a chemical or nuclear reaction, the Q of the reaction is the negative of the change in the system's total mass energy:

$$Q = M_i c^2 - M_f c^2 = -\Delta M c^2, \quad (37-50)$$

where M_i is the system's total mass before the reaction and M_f is its total mass after the reaction.

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QUESTIONS

- 1** A rod is to move at constant speed v along the x axis of reference frame S , with the rod's length parallel to that axis. An observer in frame S is to measure the length L of the rod. Which of the curves in Fig. 37-15 best gives length L (vertical axis of the graph) versus speed parameter β ?

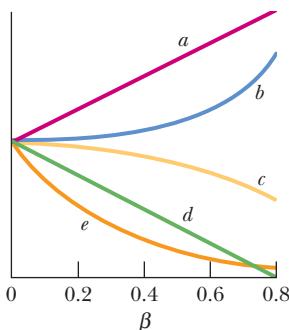


Fig. 37-15
Questions 1 and 3.

- 2** Figure 37-16 shows a ship (attached to reference frame S') passing us (standing in reference frame S). A proton is fired at nearly the speed of light along the length of the ship, from the front to the rear. (a) Is the spatial separation $\Delta x'$ between the point at which the proton is fired and the point at which it hits the ship's rear wall a positive or negative quantity? (b) Is the temporal separation $\Delta t'$ between those events a positive or negative quantity?

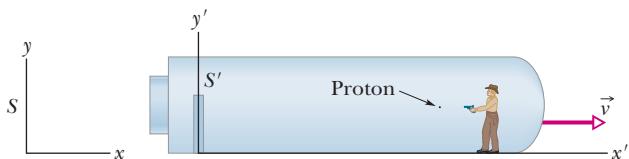


Fig. 37-16 Question 2 and Problem 68.

- 3** Reference frame S' is to pass reference frame S at speed v along the common direction of the x' and x axes, as in Fig. 37-9. An observer who rides along with frame S' is to count off 25 s on his wristwatch. The corresponding time interval Δt is to be measured by an observer in frame S . Which of the curves in Fig. 37-15 best gives Δt (vertical axis of the graph) versus speed parameter β ?

- 4** Figure 37-17 shows two clocks in stationary frame S' (they are synchronized in that frame) and one clock in moving frame S . Clocks C_1 and C'_1 read zero when they pass each other. When clocks C_1 and C'_2 pass each other, (a) which clock has the smaller reading and (b) which clock measures a proper time?

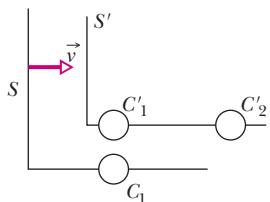


Fig. 37-17 Question 4.

- 5** Figure 37-18 shows two clocks in stationary frame S (they are synchronized in that frame) and one clock in moving frame S' . Clocks C_1 and C'_1 read zero when they pass each other. When clocks C'_1 and C_2 pass each other, (a) which clock has the smaller reading and (b) which clock measures a proper time?

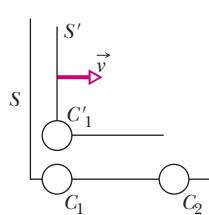


Fig. 37-18
Question 5.

- 6** Sam leaves Venus in a spaceship headed to Mars and passes Sally, who is on Earth, with a relative speed of $0.5c$. (a) Each measures the Venus–Mars voyage time. Who measures a proper time: Sam, Sally, or neither? (b) On the way, Sam sends a pulse of light to Mars. Each measures the travel time of the pulse. Who measures a proper time: Sam, Sally, or neither?

- 7** The plane of clocks and measuring rods in Fig. 37-19 is like that in Fig. 37-3. The clocks along the x axis are separated (center to cen-

- ter) by 1 light-second, as are the clocks along the y axis, and all the clocks are synchronized via the procedure described in Section 37-3. When the initial synchronizing signal of $t = 0$ from the origin reaches (a) clock A , (b) clock B , and (c) clock C , what initial time is then set on those clocks? An event occurs at clock A when it reads 10 s. (d) How long does the signal of that event take to travel to an observer stationed at the origin? (e) What time does that observer assign to the event?

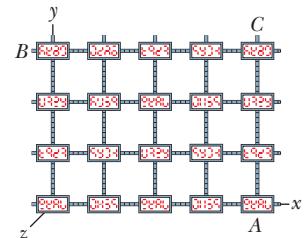


Fig. 37-19 Question 7.

- 8** The rest energy and total energy, respectively, of three particles, expressed in terms of a basic amount A are (1) A , $2A$; (2) A , $3A$; (3) $3A$, $4A$. Without written calculation, rank the particles according to their (a) mass, (b) kinetic energy, (c) Lorentz factor, and (d) speed, greatest first.

- 9** Figure 37-20 shows the triangle of Fig. 37-14 for six particles; the slanted lines 2 and 4 have the same length. Rank the particles according to (a) mass, (b) momentum magnitude, and (c) Lorentz factor, greatest first. (d) Identify which two particles have the same total energy. (e) Rank the three lowest-mass particles according to kinetic energy, greatest first.

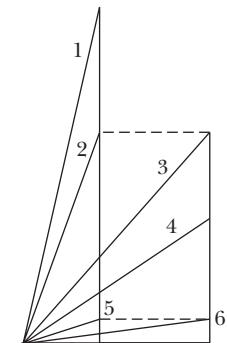


Fig. 37-20
Question 9.

- 10** While on board a starship, you intercept signals from four shuttle craft that are moving either directly toward or directly away from you. The signals have the same proper frequency f_0 . The speed and direction (both relative to you) of the shuttle craft are (a) $0.3c$ toward, (b) $0.6c$ toward, (c) $0.3c$ away, and (d) $0.6c$ away. Rank the shuttle craft according to the frequency you receive, greatest first.

- 11** Figure 37-21 shows one of four star cruisers that are in a race. As each cruiser passes the starting line, a shuttle craft leaves the cruiser and races toward the finish line. You, judging the race, are stationary relative to the starting and finish lines. The speeds v_c of the cruisers relative to you and the speeds v_s of the shuttle craft relative to their respective starships are, in that order, (1) $0.70c$, $0.40c$; (2) $0.40c$, $0.70c$; (3) $0.20c$, $0.90c$; (4) $0.50c$, $0.60c$. (a) Rank the shuttle craft according to their speeds relative to you, greatest first. (b) Rank the shuttle craft according to the distances their pilots measure from the starting line to the finish line, greatest first. (c) Each starship sends a signal to its shuttle craft at a certain frequency f_0 as measured on board the starship. Rank the shuttle craft according to the frequencies they detect, greatest first.

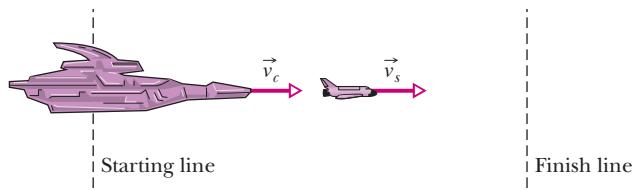


Fig. 37-21 Question 11.

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Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>

sec. 37-5 The Relativity of Time

•1 The mean lifetime of stationary muons is measured to be 2.2000 μs . The mean lifetime of high-speed muons in a burst of cosmic rays observed from Earth is measured to be 16.000 μs . To five significant figures, what is the speed parameter β of these cosmic-ray muons relative to Earth?

•2 To eight significant figures, what is speed parameter β if the Lorentz factor γ is (a) 1.010 000 0, (b) 10.000 000, (c) 100.000 00, and (d) 1000.000 0?

•3 You wish to make a round trip from Earth in a spaceship, traveling at constant speed in a straight line for exactly 6 months (as you measure the time interval) and then returning at the same constant speed. You wish further, on your return, to find Earth as it will be exactly 1000 years in the future. (a) To eight significant figures, at what speed parameter β must you travel? (b) Does it matter whether you travel in a straight line on your journey?

•4 (*Come*) back to the future. Suppose that a father is 20.00 y older than his daughter. He wants to travel outward from Earth for 2.000 y and then back to Earth for another 2.000 y (both intervals as he measures them) such that he is then 20.00 y *younger* than his daughter. What constant speed parameter β (relative to Earth) is required for the trip?

•5 **ILW** An unstable high-energy particle enters a detector and leaves a track of length 1.05 mm before it decays. Its speed relative to the detector was 0.992c. What is its proper lifetime? That is, how long would the particle have lasted before decay had it been at rest with respect to the detector?

•6 Reference frame S' is to pass reference frame S at speed v along the common direction of the x' and x axes, as in Fig. 37-9. An observer who rides along with frame S' is to count off a certain time interval on his wristwatch. The corresponding time interval Δt is to be measured by an observer in frame S . Figure 37-22 gives Δt versus speed parameter β for a range of values for β . The vertical axis scale is set by $\Delta t_a = 14.0$ s. What is interval Δt if $v = 0.98c$?

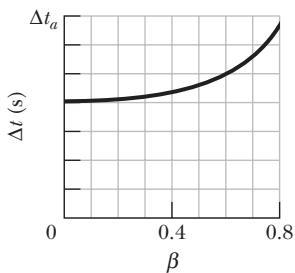


Fig. 37-22 Problem 6.

•7 The premise of the *Planet of the Apes* movies and book is that hibernating astronauts travel far into Earth's future, to a time when human civilization has been replaced by an ape civilization. Considering only special relativity, determine how far into Earth's

future the astronauts would travel if they slept for 120 y while traveling relative to Earth with a speed of 0.9990c, first outward from Earth and then back again.

sec. 37-6 The Relativity of Length

•8 An electron of $\beta = 0.999\ 987$ moves along the axis of an evacuated tube that has a length of 3.00 m as measured by a laboratory observer S at rest relative to the tube. An observer S' who is at rest relative to the electron, however, would see this tube moving with speed $v (= \beta c)$. What length would observer S' measure for the tube?

•9 **SSM** A spaceship of rest length 130 m races past a timing station at a speed of 0.740c. (a) What is the length of the spaceship as measured by the timing station? (b) What time interval will the station clock record between the passage of the front and back ends of the ship?

•10 A meter stick in frame S' makes an angle of 30° with the x' axis. If that frame moves parallel to the x axis of frame S with speed 0.90c relative to frame S , what is the length of the stick as measured from S ?

•11 A rod lies parallel to the x axis of reference frame S , moving along this axis at a speed of 0.630c. Its rest length is 1.70 m. What will be its measured length in frame S' ?

•12 The length of a spaceship is measured to be exactly half its rest length. (a) To three significant figures, what is the speed parameter β of the spaceship relative to the observer's frame? (b) By what factor do the spaceship's clocks run slow relative to clocks in the observer's frame?

•13 **ILW** A space traveler takes off from Earth and moves at speed 0.9900c toward the star Vega, which is 26.00 ly distant. How much time will have elapsed by Earth clocks (a) when the traveler reaches Vega and (b) when Earth observers receive word from the traveler that she has arrived? (c) How much older will Earth observers calculate the traveler to be (measured from her frame) when she reaches Vega than she was when she started the trip?

•14 A rod is to move at constant speed v along the x axis of reference frame S , with the rod's length parallel to that axis. An observer in frame S is to measure the length L of the rod. Figure 37-23 gives length L versus speed parameter β for a range of values for β . The vertical axis scale is set by $L_a = 1.00$ m. What is L if $v = 0.95c$?

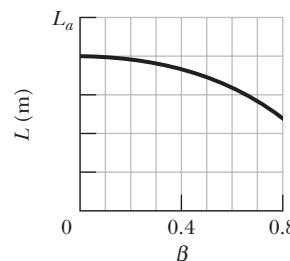


Fig. 37-23 Problem 14.

- 15** The center of our Milky Way galaxy is about 23 000 ly away. (a) To eight significant figures, at what constant speed parameter would you need to travel exactly 23 000 ly (measured in the Galaxy frame) in exactly 30 y (measured in your frame)? (b) Measured in your frame and in light-years, what length of the Galaxy would pass by you during the trip?

sec. 37-8 Some Consequences of the Lorentz Equations

- 16** Observer S reports that an event occurred on the x axis of his reference frame at $x = 3.00 \times 10^8$ m at time $t = 2.50$ s. Observer S' and her frame are moving in the positive direction of the x axis at a speed of $0.400c$. Further, $x = x' = 0$ at $t = t' = 0$. What are the (a) spatial and (b) temporal coordinate of the event according to S' ? If S' were, instead, moving in the *negative* direction of the x axis, what would be the (c) spatial and (d) temporal coordinate of the event according to S' ?

- 17 SSM WWW** In Fig. 37-9, the origins of the two frames coincide at $t = t' = 0$ and the relative speed is $0.950c$. Two micrometeorites collide at coordinates $x = 100$ km and $t = 200 \mu\text{s}$ according to an observer in frame S . What are the (a) spatial and (b) temporal coordinate of the collision according to an observer in frame S' ?

- 18** Inertial frame S' moves at a speed of $0.60c$ with respect to frame S (Fig. 37-9). Further, $x = x' = 0$ at $t = t' = 0$. Two events are recorded. In frame S , event 1 occurs at the origin at $t = 0$ and event 2 occurs on the x axis at $x = 3.0$ km at $t = 4.0 \mu\text{s}$. According to observer S' , what is the time of (a) event 1 and (b) event 2? (c) Do the two observers see the two events in the same sequence or the reverse sequence?

- 19** An experimenter arranges to trigger two flashbulbs simultaneously, producing a big flash located at the origin of his reference frame and a small flash at $x = 30.0$ km. An observer moving at a speed of $0.250c$ in the positive direction of x also views the flashes. (a) What is the time interval between them according to her? (b) Which flash does she say occurs first?

- 20** As in Fig. 37-9, reference frame S' passes reference frame S with a certain velocity. Events 1 and 2 are to have a certain temporal separation $\Delta t'$ according to the S' observer. However, their spatial separation $\Delta x'$ according to that observer has not been set yet. Figure 37-24 gives their temporal separation Δt according to the S observer as a function of $\Delta x'$ for a range of $\Delta x'$ values. The vertical axis scale is set by $\Delta t_a = 6.00 \mu\text{s}$. What is $\Delta t'$?

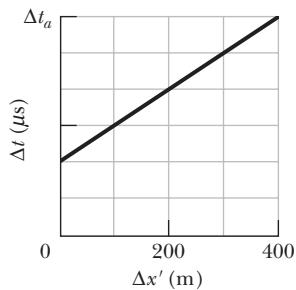


Fig. 37-24 Problem 20.

- 21 Relativistic reversal of events.** Figures 37-25a and b show the (usual) situation in which a primed reference frame passes an unprimed reference frame, in the common positive direction of the x

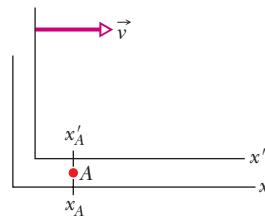
and x' axes, at a constant relative velocity of magnitude v . We are at rest in the unprimed frame; Bullwinkle, an astute student of relativity in spite of his cartoon upbringing, is at rest in the primed frame. The figures also indicate events A and B that occur at the following spacetime coordinates as measured in our unprimed frame and in Bullwinkle's primed frame:

Event	Unprimed	Primed
A	(x_A, t_A)	(x'_A, t'_A)
B	(x_B, t_B)	(x'_B, t'_B)

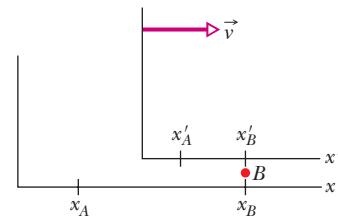
In our frame, event A occurs before event B , with temporal separation $\Delta t = t_B - t_A = 1.00 \mu\text{s}$ and spatial separation $\Delta x = x_B - x_A = 400$ m. Let $\Delta t'$ be the temporal separation of the events according to Bullwinkle. (a) Find an expression for $\Delta t'$ in terms of the speed parameter $\beta (= v/c)$ and the given data. Graph $\Delta t'$ versus β for the following two ranges of β :

- (b) 0 to 0.01 (v is low, from 0 to $0.01c$)
 (c) 0.1 to 1 (v is high, from $0.1c$ to the limit c)

- (d) At what value of β is $\Delta t' = 0$? For what range of β is the sequence of events A and B according to Bullwinkle (e) the same as ours and (f) the reverse of ours? (g) Can event A cause event B , or vice versa? Explain.



(a) Event A



(b) Event B

Fig. 37-25 Problems 21, 22, 60, and 61.

- 22** For the passing reference frames in Fig. 37-25, events A and B occur at the following spacetime coordinates: according to the unprimed frame, (x_A, t_A) and (x_B, t_B) ; according to the primed frame, (x'_A, t'_A) and (x'_B, t'_B) . In the unprimed frame, $\Delta t = t_B - t_A = 1.00 \mu\text{s}$ and $\Delta x = x_B - x_A = 400$ m. (a) Find an expression for $\Delta x'$ in terms of the speed parameter β and the given data. Graph $\Delta x'$ versus β for two ranges of β : (b) 0 to 0.01 and (c) 0.1 to 1. (d) At what value of β is $\Delta x'$ minimum, and (e) what is that minimum?

- 23 ILW** A clock moves along an x axis at a speed of $0.600c$ and reads zero as it passes the origin. (a) Calculate the clock's Lorentz factor. (b) What time does the clock read as it passes $x = 180$ m?

- 24** Bullwinkle in reference frame S' passes you in reference frame S along the common direction of the x' and x axes, as in Fig. 37-9. He carries three meter sticks: meter stick 1 is parallel to the x' axis, meter stick 2 is parallel to the y' axis, and meter stick 3 is parallel to the z' axis. On his wristwatch he counts off 15.0 s, which takes 30.0 s according to you. Two events occur during his passage. According to you, event 1 occurs at $x_1 = 33.0$ m and $t_1 = 22.0$ ns, and event 2 occurs at $x_2 = 53.0$ m and $t_2 = 62.0$ ns. According to your measurements, what is the length of (a) meter stick 1, (b) meter stick 2, and (c) meter stick 3? According to Bullwinkle, what are (d) the spatial separation and (e) the temporal separation between events 1 and 2, and (f) which event occurs first?

1052 CHAPTER 37 RELATIVITY

•25 In Fig. 37-9, observer S detects two flashes of light. A big flash occurs at $x_1 = 1200$ m and, 5.00 μs later, a small flash occurs at $x_2 = 480$ m. As detected by observer S' , the two flashes occur at a single coordinate x' . (a) What is the speed parameter of S' , and (b) is S' moving in the positive or negative direction of the x axis? To S' , (c) which flash occurs first and (d) what is the time interval between the flashes?

•26 In Fig. 37-9, observer S detects two flashes of light. A big flash occurs at $x_1 = 1200$ m and, slightly later, a small flash occurs at $x_2 = 480$ m. The time interval between the flashes is $\Delta t = t_2 - t_1$. What is the smallest value of Δt for which observer S' will determine that the two flashes occur at the same x' coordinate?

sec. 37-9 The Relativity of Velocities

•27 SSM A particle moves along the x' axis of frame S' with velocity $0.40c$. Frame S' moves with velocity $0.60c$ with respect to frame S . What is the velocity of the particle with respect to frame S ?

•28 In Fig. 37-11, frame S' moves relative to frame S with velocity $0.62\hat{c}$ while a particle moves parallel to the common x and x' axes. An observer attached to frame S' measures the particle's velocity to be $0.47\hat{c}$. In terms of c , what is the particle's velocity as measured by an observer attached to frame S according to the (a) relativistic and (b) classical velocity transformation? Suppose, instead, that the S' measure of the particle's velocity is $-0.47\hat{c}$. What velocity does the observer in S now measure according to the (c) relativistic and (d) classical velocity transformation?

•29 Galaxy A is reported to be receding from us with a speed of $0.35c$. Galaxy B, located in precisely the opposite direction, is also found to be receding from us at this same speed. What multiple of c gives the recessional speed an observer on Galaxy A would find for (a) our galaxy and (b) Galaxy B?

•30 Stellar system Q_1 moves away from us at a speed of $0.800c$. Stellar system Q_2 , which lies in the same direction in space but is closer to us, moves away from us at speed $0.400c$. What multiple of c gives the speed of Q_2 as measured by an observer in the reference frame of Q_1 ?

•31 SSM ILW WWW A spaceship whose rest length is 350 m has a speed of $0.82c$ with respect to a certain reference frame. A micrometeorite, also with a speed of $0.82c$ in this frame, passes the spaceship on an antiparallel track. How long does it take this object to pass the ship as measured on the ship?

•32 In Fig. 37-26a, particle P is to move parallel to the x and x' axes of reference frames S and S' , at a certain velocity relative to

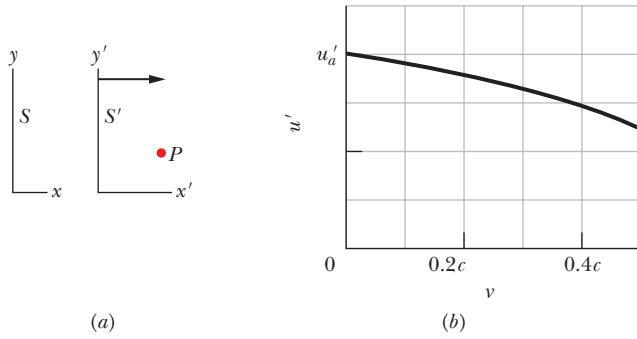


Fig. 37-26 Problem 32.

frame S . Frame S' is to move parallel to the x axis of frame S at velocity v . Figure 37-26b gives the velocity u' of the particle relative to frame S' for a range of values for v . The vertical axis scale is set by $u'_a = 0.800c$. What value will u' have if (a) $v = 0.90c$ and (b) $v \rightarrow c$?

•33 An armada of spaceships that is 1.00 ly long (in its rest frame) moves with speed $0.800c$ relative to a ground station in frame S . A messenger travels from the rear of the armada to the front with a speed of $0.950c$ relative to S . How long does the trip take as measured (a) in the messenger's rest frame, (b) in the armada's rest frame, and (c) by an observer in frame S ?

sec. 37-10 Doppler Effect for Light

•34 A sodium light source moves in a horizontal circle at a constant speed of $0.100c$ while emitting light at the proper wavelength of $\lambda_0 = 589.00$ nm. Wavelength λ is measured for that light by a detector fixed at the center of the circle. What is the wavelength shift $\lambda - \lambda_0$?

•35 SSM A spaceship, moving away from Earth at a speed of $0.900c$, reports back by transmitting at a frequency (measured in the spaceship frame) of 100 MHz. To what frequency must Earth receivers be tuned to receive the report?

•36 Certain wavelengths in the light from a galaxy in the constellation Virgo are observed to be 0.4% longer than the corresponding light from Earth sources. (a) What is the radial speed of this galaxy with respect to Earth? (b) Is the galaxy approaching or receding from Earth?

•37 Assuming that Eq. 37-36 holds, find how fast you would have to go through a red light to have it appear green. Take 620 nm as the wavelength of red light and 540 nm as the wavelength of green light.

•38 Figure 37-27 is a graph of intensity versus wavelength for light reaching Earth from galaxy NGC 7319, which is about 3×10^8 light-years away. The most intense light is emitted by the oxygen in NGC 7319. In a laboratory that emission is at wavelength $\lambda = 513$ nm, but in the light from NGC 7319 it has been shifted to 525 nm due to the Doppler effect (all the emissions from NGC 7319 have been shifted). (a) What is the radial speed of NGC 7319 relative to Earth? (b) Is the relative motion toward or away from our planet?

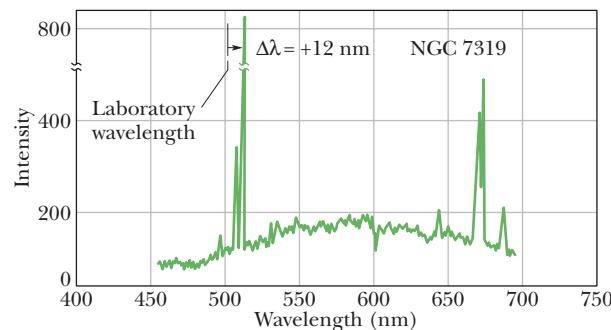


Fig. 37-27 Problem 38.

•39 SSM A spaceship is moving away from Earth at speed $0.20c$. A source on the rear of the ship emits light at wavelength 450 nm according to someone on the ship. What (a) wavelength and (b) color (blue, green, yellow, or red) are detected by someone on Earth watching the ship?

sec. 37-12 A New Look at Energy

•40 How much work must be done to increase the speed of an electron from rest to (a) $0.500c$, (b) $0.990c$, and (c) $0.9990c$?

•41 SSM WWW The mass of an electron is $9.109\ 381\ 88 \times 10^{-31}$ kg. To six significant figures, find (a) γ and (b) β for an electron with kinetic energy $K = 100.000$ MeV.

•42 What is the minimum energy that is required to break a nucleus of ^{12}C (of mass $11.996\ 71$ u) into three nuclei of ^4He (of mass $4.001\ 51$ u each)?

•43 How much work must be done to increase the speed of an electron (a) from $0.18c$ to $0.19c$ and (b) from $0.98c$ to $0.99c$? Note that the speed increase is $0.01c$ in both cases.

•44 In the reaction $\text{p} + ^{19}\text{F} \rightarrow \alpha + ^{16}\text{O}$, the masses are

$$\begin{aligned} m(\text{p}) &= 1.007825 \text{ u}, & m(\alpha) &= 4.002603 \text{ u}, \\ m(\text{F}) &= 18.998405 \text{ u}, & m(\text{O}) &= 15.994915 \text{ u}. \end{aligned}$$

Calculate the Q of the reaction from these data.

•45 In a high-energy collision between a cosmic-ray particle and a particle near the top of Earth's atmosphere, 120 km above sea level, a pion is created. The pion has a total energy E of 1.35×10^5 MeV and is traveling vertically downward. In the pion's rest frame, the pion decays 35.0 ns after its creation. At what altitude above sea level, as measured from Earth's reference frame, does the decay occur? The rest energy of a pion is 139.6 MeV.

•46 (a) If m is a particle's mass, p is its momentum magnitude, and K is its kinetic energy, show that

$$m = \frac{(pc)^2 - K^2}{2Kc^2}.$$

(b) For low particle speeds, show that the right side of the equation reduces to m . (c) If a particle has $K = 55.0$ MeV when $p = 121$ MeV/c, what is the ratio m/m_e of its mass to the electron mass?

•47 SSM A 5.00-grain aspirin tablet has a mass of 320 mg. For how many kilometers would the energy equivalent of this mass power an automobile? Assume 12.75 km/L and a heat of combustion of 3.65×10^7 J/L for the gasoline used in the automobile.

•48 The mass of a muon is 207 times the electron mass; the average lifetime of muons at rest is $2.20\ \mu\text{s}$. In a certain experiment, muons moving through a laboratory are measured to have an average lifetime of $6.90\ \mu\text{s}$. For the moving muons, what are (a) β , (b) K , and (c) p (in MeV/c)?

•49 As you read this page (on paper or monitor screen), a cosmic ray proton passes along the left-right width of the page with relative speed v and a total energy of 14.24 nJ. According to your measurements, that left-right width is 21.0 cm. (a) What is the width according to the proton's reference frame? How much time did the passage take according to (b) your frame and (c) the proton's frame?

•50 To four significant figures, find the following when the kinetic energy is 10.00 MeV: (a) γ and (b) β for an electron ($E_0 = 0.510\ 998$ MeV), (c) γ and (d) β for a proton ($E_0 = 938.272$ MeV), and (e) γ and (f) β for an α particle ($E_0 = 3727.40$ MeV).

•51 ILW What must be the momentum of a particle with mass m so that the total energy of the particle is 3.00 times its rest energy?

•52 Apply the binomial theorem (Appendix E) to the last part of Eq. 37-52 for the kinetic energy of a particle. (a) Retain the first two terms of the expansion to show the kinetic energy in the form

$$K = (\text{first term}) + (\text{second term}).$$

The first term is the classical expression for kinetic energy. The second term is the first-order correction to the classical expression. Assume the particle is an electron. If its speed v is $c/20$, what is the value of (b) the classical expression and (c) the first-order correction? If the electron's speed is $0.80c$, what is the value of (d) the classical expression and (e) the first-order correction? (f) At what speed parameter β does the first-order correction become 10% or greater of the classical expression?

•53 In Section 28-6, we showed that a particle of charge q and mass m will move in a circle of radius $r = mv/|q|B$ when its velocity \vec{v} is perpendicular to a uniform magnetic field \vec{B} . We also found that the period T of the motion is independent of speed v . These two results are approximately correct if $v \ll c$. For relativistic speeds, we must use the correct equation for the radius:

$$r = \frac{p}{|q|B} = \frac{\gamma mv}{|q|B}.$$

(a) Using this equation and the definition of period ($T = 2\pi r/v$), find the correct expression for the period. (b) Is T independent of v ? If a 10.0 MeV electron moves in a circular path in a uniform magnetic field of magnitude 2.20 T, what are (c) the radius according to Chapter 28, (d) the correct radius, (e) the period according to Chapter 28, and (f) the correct period?

•54 What is β for a particle with (a) $K = 2.00E_0$ and (b) $E = 2.00E_0$?

•55 A certain particle of mass m has momentum of magnitude mc . What are (a) β , (b) γ , and (c) the ratio K/E_0 ?

•56 (a) The energy released in the explosion of 1.00 mol of TNT is 3.40 MJ. The molar mass of TNT is 0.227 kg/mol. What weight of TNT is needed for an explosive release of 1.80×10^{14} J? (b) Can you carry that weight in a backpack, or is a truck or train required? (c) Suppose that in an explosion of a fission bomb, 0.080% of the fissionable mass is converted to released energy. What weight of fissionable material is needed for an explosive release of 1.80×10^{14} J? (d) Can you carry that weight in a backpack, or is a truck or train required?

•57 Quasars are thought to be the nuclei of active galaxies in the early stages of their formation. A typical quasar radiates energy at the rate of 10^{41} W. At what rate is the mass of this quasar being reduced to supply this energy? Express your answer in solar mass units per year, where one solar mass unit ($1\text{ smu} = 2.0 \times 10^{30}$ kg) is the mass of our Sun.

•58 The mass of an electron is $9.109\ 381\ 88 \times 10^{-31}$ kg. To eight significant figures, find the following for the given electron kinetic energy: (a) γ and (b) β for $K = 1.000\ 000\ 0$ keV, (c) γ and (d) β for $K = 1.000\ 000\ 0$ MeV, and then (e) γ and (f) β for $K = 1.000\ 000\ 0$ GeV.

•59 An alpha particle with kinetic energy 7.70 MeV collides with an ^{14}N nucleus at rest, and the two transform into an ^{17}O nucleus and a proton. The proton is emitted at 90° to the direction of the incident alpha particle and has a kinetic energy of 4.44 MeV. The masses of the various particles are alpha particle, 4.00260 u; ^{14}N , 14.00307 u; proton, 1.007825 u; and ^{17}O , 16.99914 u.

1054 CHAPTER 37 RELATIVITY

In MeV, what are (a) the kinetic energy of the oxygen nucleus and (b) the Q of the reaction? (*Hint:* The speeds of the particles are much less than c .)

Additional Problems

60 *Temporal separation between two events.* Events A and B occur with the following spacetime coordinates in the reference frames of Fig. 37-25: according to the unprimed frame, (x_A, t_A) and (x_B, t_B) ; according to the primed frame, (x'_A, t'_A) and (x'_B, t'_B) . In the unprimed frame, $\Delta t = t_B - t_A = 1.00 \mu\text{s}$ and $\Delta x = x_B - x_A = 240 \text{ m}$. (a) Find an expression for $\Delta t'$ in terms of the speed parameter β and the given data. Graph $\Delta t'$ versus β for the following two ranges of β : (b) 0 to 0.01 and (c) 0.1 to 1. (d) At what value of β is $\Delta t'$ minimum and (e) what is that minimum? (f) Can one of these events cause the other? Explain.

61 *Spatial separation between two events.* For the passing reference frames of Fig. 37-25, events A and B occur with the following spacetime coordinates: according to the unprimed frame, (x_A, t_A) and (x_B, t_B) ; according to the primed frame, (x'_A, t'_A) and (x'_B, t'_B) . In the unprimed frame, $\Delta t = t_B - t_A = 1.00 \mu\text{s}$ and $\Delta x = x_B - x_A = 240 \text{ m}$. (a) Find an expression for $\Delta x'$ in terms of the speed parameter β and the given data. Graph $\Delta x'$ versus β for two ranges of β : (b) 0 to 0.01 and (c) 0.1 to 1. (d) At what value of β is $\Delta x' = 0$?

62 In Fig. 37-28a, particle P is to move parallel to the x and x' axes of reference frames S and S' , at a certain velocity relative to frame S . Frame S' is to move parallel to the x axis of frame S at velocity v . Figure 37-28b gives the velocity u' of the particle relative to frame S' for a range of values for v . The vertical axis scale is set by $u'_a = -0.800c$. What value will u' have if (a) $v = 0.80c$ and (b) $v \rightarrow c$?

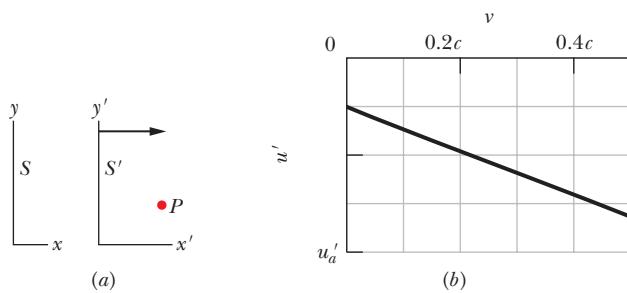


Fig. 37-28 Problem 62.

63 *Superluminal jets.* Figure 37-29a shows the path taken by a knot in a jet of ionized gas that has been expelled from a galaxy. The knot travels at constant velocity \vec{v} at angle θ from the direction of Earth. The knot occasionally emits a burst of light, which is eventually detected on Earth. Two bursts are indicated in Fig. 37-29a, separated by time t as measured in a stationary frame near the bursts. The bursts are shown in Fig. 37-29b as if they were photographed on the same piece of film, first when light from burst 1 arrived on Earth and then later when light from burst 2 arrived. The apparent distance D_{app} traveled by the knot between the two bursts is the distance across an Earth-observer's view of the knot's path. The apparent time T_{app} between the bursts is the difference in the arrival times of the light from them. The apparent speed of the knot is then $V_{\text{app}} = D_{\text{app}}/T_{\text{app}}$. In terms of v , t , and θ , what are (a) D_{app} and (b) T_{app} ? (c) Evaluate V_{app} for $v = 0.980c$ and $\theta = 30.0^\circ$.

When superluminal (faster than light) jets were first observed, they seemed to defy special relativity—at least until the correct geometry (Fig. 37-29a) was understood.

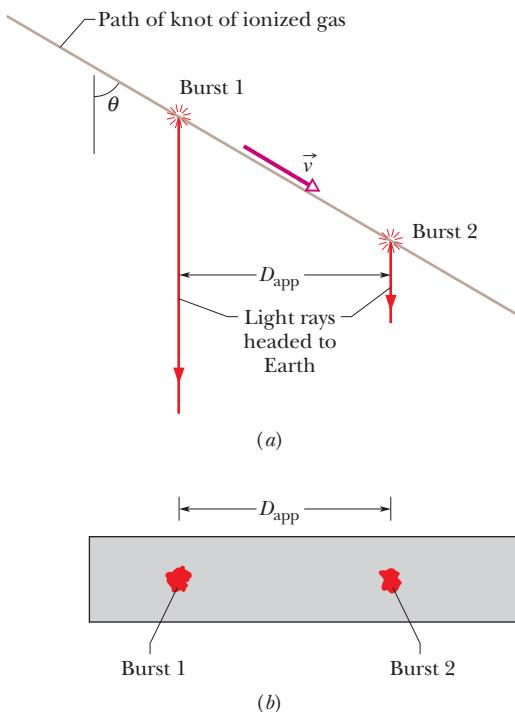


Fig. 37-29 Problem 63.

64 Reference frame S' passes reference frame S with a certain velocity as in Fig. 37-9. Events 1 and 2 are to have a certain spatial separation $\Delta x'$ according to the S' observer. However, their temporal separation $\Delta t'$ according to that observer has not been set yet. Figure 37-30 gives their spatial separation Δx according to the S observer as a function of $\Delta t'$ for a range of $\Delta t'$ values. The vertical axis scale is set by $\Delta x_a = 10.0 \text{ m}$. What is $\Delta x'$?

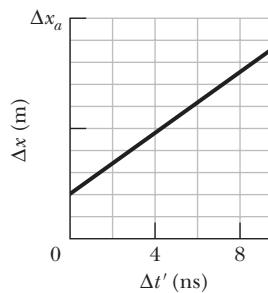


Fig. 37-30 Problem 64.

65 *Another approach to velocity transformations.* In Fig. 37-31, reference frames B and C move past reference frame A in the com-

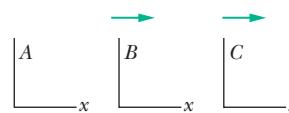


Fig. 37-31 Problems 65, 66, and 67.

mon direction of their x axes. Represent the x components of the velocities of one frame relative to another with a two-letter subscript. For example, v_{AB} is the x component of the velocity of A relative to B . Similarly, represent the corresponding speed parameters with two-letter subscripts. For example, β_{AB} ($= v_{AB}/c$) is the speed parameter corresponding to v_{AB} . (a) Show that

$$\beta_{AC} = \frac{\beta_{AB} + \beta_{BC}}{1 + \beta_{AB}\beta_{BC}}.$$

Let M_{AB} represent the ratio $(1 - \beta_{AB})/(1 + \beta_{AB})$, and let M_{BC} and M_{AC} represent similar ratios. (b) Show that the relation

$$M_{AC} = M_{AB}M_{BC}$$

is true by deriving the equation of part (a) from it.

66 Continuation of Problem 65. Use the result of part (b) in Problem 65 for the motion along a single axis in the following situation. Frame A in Fig. 37-31 is attached to a particle that moves with velocity $+0.500c$ past frame B , which moves past frame C with a velocity of $+0.500c$. What are (a) M_{AC} , (b) β_{AC} , and (c) the velocity of the particle relative to frame C' ?

67 Continuation of Problem 65. Let reference frame C in Fig. 37-31 move past reference frame D (not shown). (a) Show that

$$M_{AD} = M_{AB}M_{BC}M_{CD}.$$

(b) Now put this general result to work: Three particles move parallel to a single axis on which an observer is stationed. Let plus and minus signs indicate the directions of motion along that axis. Particle A moves past particle B at $\beta_{AB} = +0.20$. Particle B moves past particle C at $\beta_{BC} = -0.40$. Particle C moves past observer D at $\beta_{CD} = +0.60$. What is the velocity of particle A relative to observer D ? (The solution technique here is *much* faster than using Eq. 37-29.)

68 Figure 37-16 shows a ship (attached to reference frame S') passing us (standing in reference frame S) with velocity $\vec{v} = 0.950\hat{c}$. A proton is fired at speed $0.980c$ relative to the ship from the front of the ship to the rear. The proper length of the ship is 760 m. What is the temporal separation between the time the proton is fired and the time it hits the rear wall of the ship according to (a) a passenger in the ship and (b) us? Suppose that, instead, the proton is fired from the rear to the front. What then is the temporal separation between the time it is fired and the time it hits the front wall according to (c) the passenger and (d) us?

69 The car-in-the-garage problem. Carman has just purchased the world's longest stretch limo, which has a proper length of $L_c = 30.5$ m. In Fig. 37-32a, it is shown parked in front of a garage with a proper length of $L_g = 6.00$ m. The garage has a front door (shown open) and a back door (shown closed). The limo is obviously longer than the garage. Still, Garageman, who owns the garage and knows something about relativistic length contraction, makes a bet with Carman that the limo can fit in the garage with both doors closed. Carman, who dropped his physics course before reaching special relativity, says such a thing, even in principle, is impossible.

To analyze Garageman's scheme, an x_c axis is attached to the limo, with $x_c = 0$ at the rear bumper, and an x_g axis is attached to the garage, with $x_g = 0$ at the (now open) front door. Then Carman is to drive the limo directly toward the front door at a velocity of $0.9980c$ (which is, of course, both technically and financially impossible). Carman is stationary in the x_c reference frame; Garageman is stationary in the x_g reference frame.

There are two events to consider. *Event 1:* When the rear bumper clears the front door, the front door is closed. Let the time of this event be zero to both Carman and Garageman: $t_{g1} = t_{c1} = 0$. The event occurs at $x_c = x_g = 0$. Figure 37-32b shows event 1 according to the x_g reference frame. *Event 2:* When the front bumper reaches the back door, that door opens. Figure 37-32c shows event 2 according to the x_g reference frame.

According to Garageman, (a) what is the length of the limo, and what are the spacetime coordinates (b) x_{g2} and (c) t_{g2} of event 2? (d) For how long is the limo temporarily "trapped" inside the garage with both doors shut? Now consider the situation from the x_c reference frame, in which the garage comes racing past the limo at a velocity of $-0.9980c$. According to Carman, (e) what is the length of the passing garage, what are the spacetime coordinates (f) x_{c2} and (g) t_{c2} of event 2, (h) is the limo ever in the garage with both doors shut, and (i) which event occurs first? (j) Sketch events 1 and 2 as seen by Carman. (k) Are the events causally related; that is, does one of them cause the other? (l) Finally, who wins the bet?

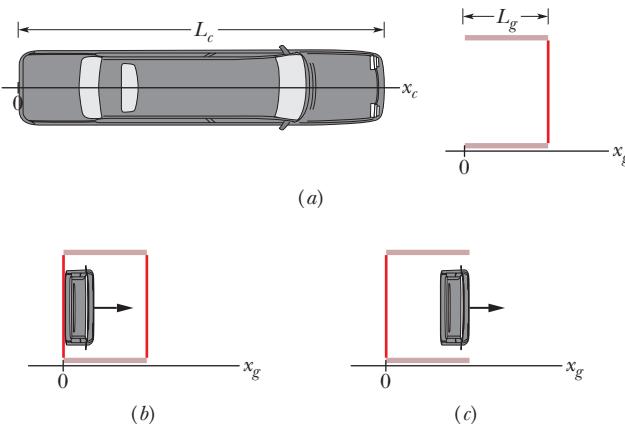


Fig. 37-32 Problem 69.

70 An airplane whose rest length is 40.0 m is moving at uniform velocity with respect to Earth, at a speed of 630 m/s. (a) By what fraction of its rest length is it shortened to an observer on Earth? (b) How long would it take, according to Earth clocks, for the airplane's clock to fall behind by $1.00\ \mu s$?

71 SSM To circle Earth in low orbit, a satellite must have a speed of about 2.7×10^4 km/h. Suppose that two such satellites orbit Earth in opposite directions. (a) What is their relative speed as they pass, according to the classical Galilean velocity transformation equation? (b) What fractional error do you make in (a) by not using the (correct) relativistic transformation equation?

72 Find the speed parameter of a particle that takes 2.0 y longer than light to travel a distance of 6.0 ly.

73 SSM How much work is needed to accelerate a proton from a speed of $0.9850c$ to a speed of $0.9860c$?

74 A pion is created in the higher reaches of Earth's atmosphere when an incoming high-energy cosmic-ray particle collides with an atomic nucleus. A pion so formed descends toward Earth with a speed of $0.99c$. In a reference frame in which they are at rest, pions decay with an average life of 26 ns. As measured in a frame fixed with respect to Earth, how far (on the average) will such a pion move through the atmosphere before it decays?

1056 CHAPTER 37 RELATIVITY

75 SSM If we intercept an electron having total energy 1533 MeV that came from Vega, which is 26 ly from us, how far in light-years was the trip in the rest frame of the electron?

76 The total energy of a proton passing through a laboratory apparatus is 10.611 nJ. What is its speed parameter β ? Use the proton mass given in Appendix B under “Best Value,” not the commonly remembered rounded number.

77 A spaceship at rest in a certain reference frame S is given a speed increment of $0.50c$. Relative to its new rest frame, it is then given a further $0.50c$ increment. This process is continued until its speed with respect to its original frame S exceeds $0.999c$. How many increments does this process require?

78 In the red shift of radiation from a distant galaxy, a certain radiation, known to have a wavelength of 434 nm when observed in the laboratory, has a wavelength of 462 nm. (a) What is the radial speed of the galaxy relative to Earth? (b) Is the galaxy approaching or receding from Earth?

79 SSM What is the momentum in MeV/c of an electron with a kinetic energy of 2.00 MeV?

80 The radius of Earth is 6370 km, and its orbital speed about the Sun is 30 km/s. Suppose Earth moves past an observer at this speed. To the observer, by how much does Earth’s diameter contract along the direction of motion?

81 A particle with mass m has speed $c/2$ relative to inertial frame S . The particle collides with an identical particle at rest relative to frame S . Relative to S , what is the speed of a frame S' in which the total momentum of these particles is zero? This frame is called the *center of momentum frame*.

82 An elementary particle produced in a laboratory experiment travels 0.230 mm through the lab at a relative speed of $0.960c$ before it decays (becomes another particle). (a) What is the proper lifetime of the particle? (b) What is the distance the particle travels as measured from its rest frame?

83 What are (a) K , (b) E , and (c) p (in GeV/c) for a proton moving at speed $0.990c$? What are (d) K , (e) E , and (f) p (in MeV/c) for an electron moving at speed $0.990c$?

84 A radar transmitter T is fixed to a reference frame S' that is moving to the right with speed v relative to reference frame S (Fig. 37-33). A mechanical timer (essentially a clock) in frame S' , having a period τ_0 (measured in S'), causes transmitter T to emit timed radar pulses, which travel at the speed of light and are received by R , a receiver fixed in frame S . (a) What is the period τ of the timer as detected by observer A , who is fixed in frame S ? (b) Show that at receiver R the time interval between pulses arriving from T is not τ or τ_0 , but

$$\tau_R = \tau_0 \sqrt{\frac{c + v}{c - v}}.$$

(c) Explain why receiver R and observer A , who are in the same

reference frame, measure a different period for the transmitter. (Hint: A clock and a radar pulse are not the same thing.)

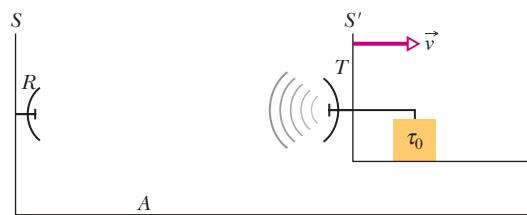


Fig. 37-33 Problem 84.

85 One cosmic-ray particle approaches Earth along Earth’s north–south axis with a speed of $0.80c$ toward the geographic north pole, and another approaches with a speed of $0.60c$ toward the geographic south pole (Fig. 37-34). What is the relative speed of approach of one particle with respect to the other?

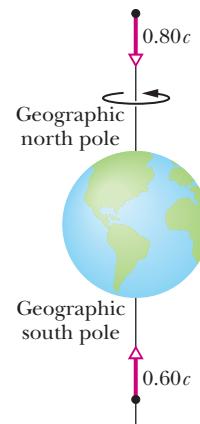


Fig. 37-34 Problem 85.

86 (a) How much energy is released in the explosion of a fission bomb containing 3.0 kg of fissionable material? Assume that 0.10% of the mass is converted to released energy. (b) What mass of TNT would have to explode to provide the same energy release? Assume that each mole of TNT liberates 3.4 MJ of energy on exploding. The molecular mass of TNT is 0.227 kg/mol. (c) For the same mass of explosive, what is the ratio of the energy released in a nuclear explosion to that released in a TNT explosion?

87 (a) What potential difference would accelerate an electron to speed c according to classical physics? (b) With this potential difference, what speed would the electron actually attain?

88 A Foron cruiser moving directly toward a Reptilian scout ship fires a decoy toward the scout ship. Relative to the scout ship, the speed of the decoy is $0.980c$ and the speed of the Foron cruiser is $0.900c$. What is the speed of the decoy relative to the cruiser?

38

PHOTONS AND MATTER WAVES

38-1 WHAT IS PHYSICS?

One primary focus of physics is Einstein's theory of relativity, which took us into a world far beyond that of ordinary experience—the world of objects moving at speeds close to the speed of light. Among other surprises, Einstein's theory predicts that the rate at which a clock runs depends on how fast the clock is moving relative to the observer: the faster the motion, the slower the clock rate. This and other predictions of the theory have passed every experimental test devised thus far, and relativity theory has led us to a deeper and more satisfying view of the nature of space and time.

Now you are about to explore a second world that is outside ordinary experience—the subatomic world. You will encounter a new set of surprises that, though they may sometimes seem bizarre, have led physicists step by step to a deeper view of reality.

Quantum physics, as our new subject is called, answers such questions as: Why do the stars shine? Why do the elements exhibit the order that is so apparent in the periodic table? How do transistors and other microelectronic devices work? Why does copper conduct electricity but glass does not? Because quantum physics accounts for all of chemistry, including biochemistry, we need to understand it if we are to understand life itself.

Some of the predictions of quantum physics seem strange even to the physicists and philosophers who study its foundations. Still, experiment after experiment has proved the theory correct, and many have exposed even stranger aspects of the theory. The quantum world is an amusement park full of wonderful rides that are guaranteed to shake up the commonsense world view you have developed since childhood. We begin our exploration of that quantum park with the photon.

38-2 The Photon, the Quantum of Light

Quantum physics (which is also known as *quantum mechanics* and *quantum theory*) is largely the study of the microscopic world. In that world, many quantities are found only in certain minimum (*elementary*) amounts, or integer multiples of those elementary amounts; these quantities are then said to be *quantized*. The elementary amount that is associated with such a quantity is called the **quantum** of that quantity (*quanta* is the plural).

In a loose sense, U.S. currency is quantized because the coin of least value is the penny, or \$0.01 coin, and the values of all other coins and bills are restricted to integer multiples of that least amount. In other words, the currency quantum is \$0.01, and all greater amounts of currency are of the form $n(\$0.01)$, where n is always a positive integer. For example, you cannot hand someone $\$0.755 = 75.5(\$0.01)$.

In 1905, Einstein proposed that electromagnetic radiation (or simply *light*) is quantized and exists in elementary amounts (quanta) that we now call **photons**. This proposal should seem strange to you because we have just spent several chapters discussing the classical idea that light is a sinusoidal wave, with a wavelength λ , a frequency f , and a speed c such that

$$f = \frac{c}{\lambda}. \quad (38-1)$$

Furthermore, in Chapter 33 we discussed the classical light wave as being an interdependent combination of electric and magnetic fields, each oscillating at frequency f . How can this wave of oscillating fields consist of an elementary amount of something—the light quantum? What *is* a photon?

The concept of a light quantum, or a photon, turns out to be far more subtle and mysterious than Einstein imagined. Indeed, it is still very poorly understood. In this book, we shall discuss only some of the basic aspects of the photon concept, somewhat along the lines of Einstein's proposal.

According to that proposal, the quantum of a light wave of frequency f has the energy

$$E = hf \quad (\text{photon energy}). \quad (38-2)$$

Here h is the **Planck constant**, the constant we first met in Eq. 32-23, and which has the value

$$h = 6.63 \times 10^{-34} \text{ J} \cdot \text{s} = 4.14 \times 10^{-15} \text{ eV} \cdot \text{s}. \quad (38-3)$$

The smallest amount of energy a light wave of frequency f can have is hf , the energy of a single photon. If the wave has more energy, its total energy must be an integer multiple of hf , just as the currency in our previous example must be an integer multiple of \$0.01. The light cannot have an energy of, say, $0.6hf$ or $75.5hf$.

Einstein further proposed that when light is absorbed or emitted by an object (matter), the absorption or emission event occurs in the atoms of the object. When light of frequency f is absorbed by an atom, the energy hf of one photon is transferred from the light to the atom. In this *absorption event*, the photon vanishes and the atom is said to absorb it. When light of frequency f is emitted by an atom, an amount of energy hf is transferred from the atom to the light. In this *emission event*, a photon suddenly appears and the atom is said to emit it. Thus, we can have *photon absorption* and *photon emission* by atoms in an object.

For an object consisting of many atoms, there can be many photon absorptions (such as with sunglasses) or photon emissions (such as with lamps). However, each absorption or emission event still involves the transfer of an amount of energy equal to that of a single photon of the light.

When we discussed the absorption or emission of light in previous chapters, our examples involved so much light that we had no need of quantum physics, and we got by with classical physics. However, in the late 20th century, technology became advanced enough that single-photon experiments could be conducted and put to practical use. Since then quantum physics has become part of standard engineering practice, especially in optical engineering.



CHECKPOINT 1

Rank the following radiations according to their associated photon energies, greatest first: (a) yellow light from a sodium vapor lamp, (b) a gamma ray emitted by a radioactive nucleus, (c) a radio wave emitted by the antenna of a commercial radio station, (d) a microwave beam emitted by airport traffic control radar.

Sample Problem

Emission and absorption of light as photons

A sodium vapor lamp is placed at the center of a large sphere that absorbs all the light reaching it. The rate at which the lamp emits energy is 100 W; assume that the emission is entirely at a wavelength of 590 nm. At what rate are photons absorbed by the sphere?

KEY IDEAS

The light is emitted and absorbed as photons. We assume that all the light emitted by the lamp reaches (and thus is absorbed by) the sphere. So, the rate R at which photons are absorbed by the sphere is equal to the rate R_{emit} at which photons are emitted by the lamp.

Calculations: That rate is

$$R_{\text{emit}} = \frac{\text{rate of energy emission}}{\text{energy per emitted photon}} = \frac{P_{\text{emit}}}{E}.$$

$$R = R_{\text{emit}} = \frac{P_{\text{emit}}}{hf}.$$

Using Eq. 38-1 ($f = c/\lambda$) to substitute for f and then entering known data, we obtain

$$\begin{aligned} R &= \frac{P_{\text{emit}}\lambda}{hc} \\ &= \frac{(100 \text{ W})(590 \times 10^{-9} \text{ m})}{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(2.998 \times 10^8 \text{ m/s})} \\ &= 2.97 \times 10^{20} \text{ photons/s.} \end{aligned} \quad (\text{Answer})$$



Additional examples, video, and practice available at WileyPLUS

38-3 The Photoelectric Effect

If you direct a beam of light of short enough wavelength onto a clean metal surface, the light will cause electrons to leave that surface (the light will *eject* the electrons from the surface). This **photoelectric effect** is used in many devices, including TV cameras, camcorders, and night vision viewers. Einstein supported his photon concept by using it to explain this effect, which simply cannot be understood without quantum physics.

Let us analyze two basic photoelectric experiments, each using the apparatus of Fig. 38-1, in which light of frequency f is directed onto target T and ejects electrons from it. A potential difference V is maintained between target T and collector cup C to sweep up these electrons, said to be **photoelectrons**. This collection produces a **photoelectric current** i that is measured with meter A.

First Photoelectric Experiment

We adjust the potential difference V by moving the sliding contact in Fig. 38-1 so that collector C is slightly negative with respect to target T. This potential difference acts to slow down the ejected electrons. We then vary V until it reaches a certain value, called the **stopping potential** V_{stop} , at which point the reading of meter A has just dropped to zero. When $V = V_{\text{stop}}$, the most energetic ejected electrons are turned back just before reaching the collector. Then K_{max} , the kinetic energy of these most energetic electrons, is

$$K_{\text{max}} = eV_{\text{stop}}, \quad (38-4)$$

where e is the elementary charge.

Measurements show that for light of a given frequency, K_{max} does not depend on the intensity of the light source. Whether the source is dazzling bright or so feeble that you can scarcely detect it (or has some intermediate brightness), the maximum kinetic energy of the ejected electrons always has the same value.

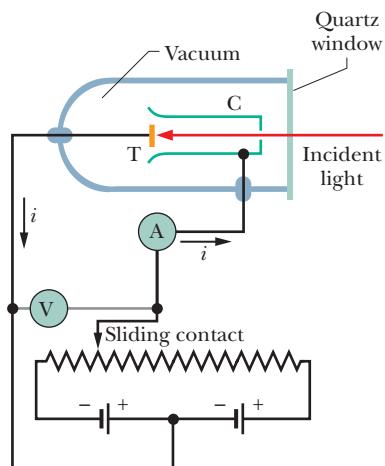


Fig. 38-1 An apparatus used to study the photoelectric effect. The incident light shines on target T, ejecting electrons, which are collected by collector cup C. The electrons move in the circuit in a direction opposite the conventional current arrows. The batteries and the variable resistor are used to produce and adjust the electric potential difference between T and C.

This experimental result is a puzzle for classical physics. Classically, the incident light is a sinusoidally oscillating electromagnetic wave. An electron in the target should oscillate sinusoidally due to the oscillating electric force on it from the wave's electric field. If the amplitude of the electron's oscillation is great enough, the electron should break free of the target's surface—that is, be ejected from the target. Thus, if we increase the amplitude of the wave and its oscillating electric field, the electron should get a more energetic “kick” as it is being ejected. *However, that is not what happens.* For a given frequency, intense light beams and feeble light beams give exactly the same maximum kick to ejected electrons.

The actual result follows naturally if we think in terms of photons. Now the energy that can be transferred from the incident light to an electron in the target is that of a single photon. Increasing the light intensity increases the *number* of photons in the light, but the photon energy, given by Eq. 38-2 ($E = hf$), is unchanged because the frequency is unchanged. Thus, the energy transferred to the kinetic energy of an electron is also unchanged.

Second Photoelectric Experiment

Now we vary the frequency f of the incident light and measure the associated stopping potential V_{stop} . Figure 38-2 is a plot of V_{stop} versus f . Note that the photoelectric effect does not occur if the frequency is below a certain **cutoff frequency** f_0 or, equivalently, if the wavelength is greater than the corresponding **cutoff wavelength** $\lambda_0 = c/f_0$. This is so *no matter how intense the incident light is*.

This is another puzzle for classical physics. If you view light as an electromagnetic wave, you must expect that no matter how low the frequency, electrons can always be ejected by light if you supply them with enough energy—that is, if you use a light source that is bright enough. *That is not what happens.* For light below the cutoff frequency f_0 , the photoelectric effect does not occur, no matter how bright the light source.

The existence of a cutoff frequency is, however, just what we should expect if the energy is transferred via photons. The electrons within the target are held there by electric forces. (If they weren't, they would drip out of the target due to the gravitational force on them.) To just escape from the target, an electron must pick up a certain minimum energy Φ , where Φ is a property of the target material called its **work function**. If the energy hf transferred to an electron by a photon exceeds the work function of the material (if $hf > \Phi$), the electron can escape the target. If the energy transferred does not exceed the work function (that is, if $hf < \Phi$), the electron cannot escape. This is what Fig. 38-2 shows.

The Photoelectric Equation

Einstein summed up the results of such photoelectric experiments in the equation

$$hf = K_{\max} + \Phi \quad (\text{photoelectric equation}). \quad (38-5)$$

This is a statement of the conservation of energy for a single photon absorption by a target with work function Φ . Energy equal to the photon's energy hf is transferred to a single electron in the material of the target. If the electron is to escape from the target, it must pick up energy at least equal to Φ . Any additional energy ($hf - \Phi$) that the electron acquires from the photon appears as kinetic energy K of the electron. In the most favorable circumstance, the electron can escape through the surface without losing any of this kinetic energy in the process; it then appears outside the target with the maximum possible kinetic energy K_{\max} .

38-3 THE PHOTOELECTRIC EFFECT

1061

Electrons can escape only if the light frequency exceeds a certain value.

The escaping electron's kinetic energy is greater for a greater light frequency.

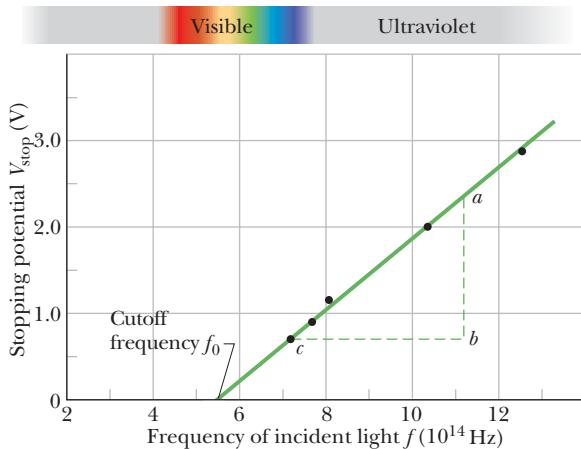


Fig. 38-2 The stopping potential V_{stop} as a function of the frequency f of the incident light for a sodium target T in the apparatus of Fig. 38-1. (Data reported by R. A. Millikan in 1916.)

Let us rewrite Eq. 38-5 by substituting for K_{\max} from Eq. 38-4 ($K_{\max} = eV_{\text{stop}}$). After a little rearranging we get

$$V_{\text{stop}} = \left(\frac{h}{e}\right)f - \frac{\Phi}{e}. \quad (38-6)$$

The ratios h/e and Φ/e are constants, and so we would expect a plot of the measured stopping potential V_{stop} versus the frequency f of the light to be a straight line, as it is in Fig. 38-2. Further, the slope of that straight line should be h/e . As a check, we measure ab and bc in Fig. 38-2 and write

$$\begin{aligned} \frac{h}{e} &= \frac{ab}{bc} = \frac{2.35 \text{ V} - 0.72 \text{ V}}{(11.2 \times 10^{14} \text{ Hz} - 7.2 \times 10^{14} \text{ Hz})} \\ &= 4.1 \times 10^{-15} \text{ V} \cdot \text{s}. \end{aligned}$$

Multiplying this result by the elementary charge e , we find

$$h = (4.1 \times 10^{-15} \text{ V} \cdot \text{s})(1.6 \times 10^{-19} \text{ C}) = 6.6 \times 10^{-34} \text{ J} \cdot \text{s},$$

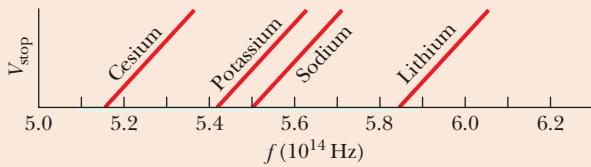
which agrees with values measured by many other methods.

An aside: An explanation of the photoelectric effect certainly requires quantum physics. For many years, Einstein's explanation was also a compelling argument for the existence of photons. However, in 1969 an alternative explanation for the effect was found that used quantum physics but did not need the concept of photons. Light *is* in fact quantized as photons, but Einstein's explanation of the photoelectric effect is not the best argument for that fact.



CHECKPOINT 2

The figure shows data like those of Fig. 38-2 for targets of cesium, potassium, sodium, and lithium. The plots are parallel. (a) Rank the targets according to their work functions, greatest first. (b) Rank the plots according to the value of h they yield, greatest first.



Sample Problem**Photoelectric effect and work function**

Find the work function Φ of sodium from Fig. 38-2.

KEY IDEAS

We can find the work function Φ from the cutoff frequency f_0 (which we can measure on the plot). The reasoning is this: At the cutoff frequency, the kinetic energy K_{\max} in Eq. 38-5 is zero. Thus, all the energy hf that is transferred from a photon to an electron goes into the electron's escape, which requires an energy of Φ .

Calculations: From that last idea, Eq. 38-5 then gives us, with $f = f_0$,

$$hf_0 = 0 + \Phi = \Phi.$$

In Fig. 38-2, the cutoff frequency f_0 is the frequency at which the plotted line intercepts the horizontal frequency axis, about 5.5×10^{14} Hz. We then have

$$\begin{aligned} \Phi &= hf_0 = (6.63 \times 10^{-34} \text{ J}\cdot\text{s})(5.5 \times 10^{14} \text{ Hz}) \\ &= 3.6 \times 10^{-19} \text{ J} = 2.3 \text{ eV}. \end{aligned} \quad (\text{Answer})$$



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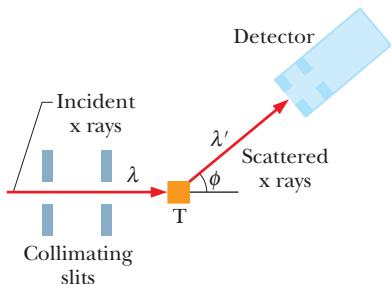


Fig. 38-3 Compton's apparatus. A beam of x rays of wavelength $\lambda = 71.1$ pm is directed onto a carbon target T. The x rays scattered from the target are observed at various angles ϕ to the direction of the incident beam. The detector measures both the intensity of the scattered x rays and their wavelength.

38-4 Photons Have Momentum

In 1916, Einstein extended his concept of light quanta (photons) by proposing that a quantum of light has linear momentum. For a photon with energy hf , the magnitude of that momentum is

$$p = \frac{hf}{c} = \frac{h}{\lambda} \quad (\text{photon momentum}), \quad (38-7)$$

where we have substituted for f from Eq. 38-1 ($f = c/\lambda$). Thus, when a photon interacts with matter, energy *and* momentum are transferred, *as if* there were a collision between the photon and matter in the classical sense (as in Chapter 9).

In 1923, Arthur Compton at Washington University in St. Louis carried out an experiment that supported the view that both momentum and energy are transferred via photons. He arranged for a beam of x rays of wavelength λ to be directed onto a target made of carbon, as shown in Fig. 38-3. An x ray is a form of electromagnetic radiation, at high frequency and thus small wavelength. Compton measured the wavelengths and intensities of the x rays that were scattered in various directions from his carbon target.

Figure 38-4 shows his results. Although there is only a single wavelength ($\lambda = 71.1$ pm) in the incident x-ray beam, we see that the scattered x rays contain a range of wavelengths with two prominent intensity peaks. One peak is centered about the incident wavelength λ , the other about a wavelength λ' that is longer than λ by an amount $\Delta\lambda$, which is called the **Compton shift**. The value

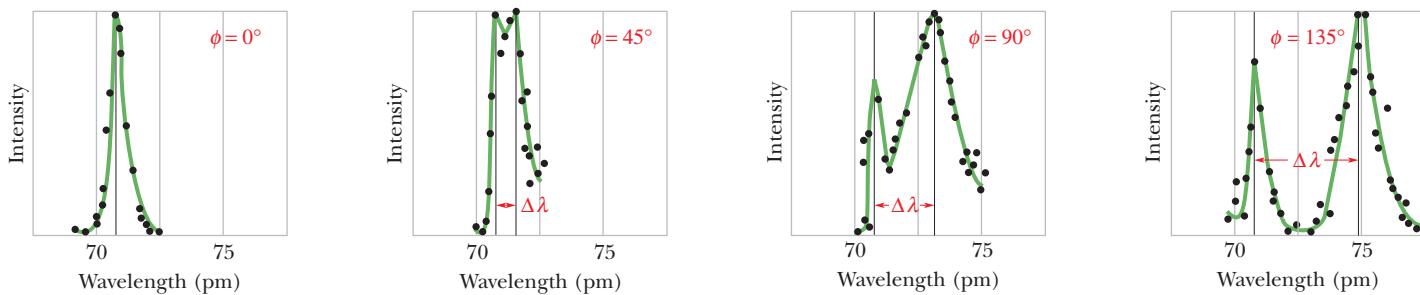


Fig. 38-4 Compton's results for four values of the scattering angle ϕ . Note that the Compton shift $\Delta\lambda$ increases as the scattering angle increases.

38-4 PHOTONS HAVE MOMENTUM

1063

of the Compton shift varies with the angle at which the scattered x rays are detected and is greater for a greater angle.

Figure 38-4 is still another puzzle for classical physics. Classically, the incident x-ray beam is a sinusoidally oscillating electromagnetic wave. An electron in the carbon target should oscillate sinusoidally due to the oscillating electric force on it from the wave's electric field. Further, the electron should oscillate at the same frequency as the wave and should send out waves *at this same frequency*, as if it were a tiny transmitting antenna. Thus, the x rays scattered by the electron should have the same frequency, and the same wavelength, as the x rays in the incident beam—but they don't.

Compton interpreted the scattering of x rays from carbon in terms of energy and momentum transfers, via photons, between the incident x-ray beam and loosely bound electrons in the carbon target. Let us see, first conceptually and then quantitatively, how this quantum physics interpretation leads to an understanding of Compton's results.

Suppose a single photon (of energy $E = hf$) is associated with the interaction between the incident x-ray beam and a stationary electron. In general, the direction of travel of the x ray will change (the x ray is scattered), and the electron will recoil, which means that the electron has obtained some kinetic energy. Energy is conserved in this isolated interaction. Thus, the energy of the scattered photon ($E' = hf'$) must be less than that of the incident photon. The scattered x rays must then have a lower frequency f' and thus a longer wavelength λ' than the incident x rays, just as Compton's experimental results in Fig. 38-4 show.

For the quantitative part, we first apply the law of conservation of energy. Figure 38-5 suggests a “collision” between an x ray and an initially stationary free electron in the target. As a result of the collision, an x ray of wavelength λ' moves off at an angle ϕ and the electron moves off at an angle θ , as shown. Conservation of energy then gives us

$$hf = hf' + K,$$

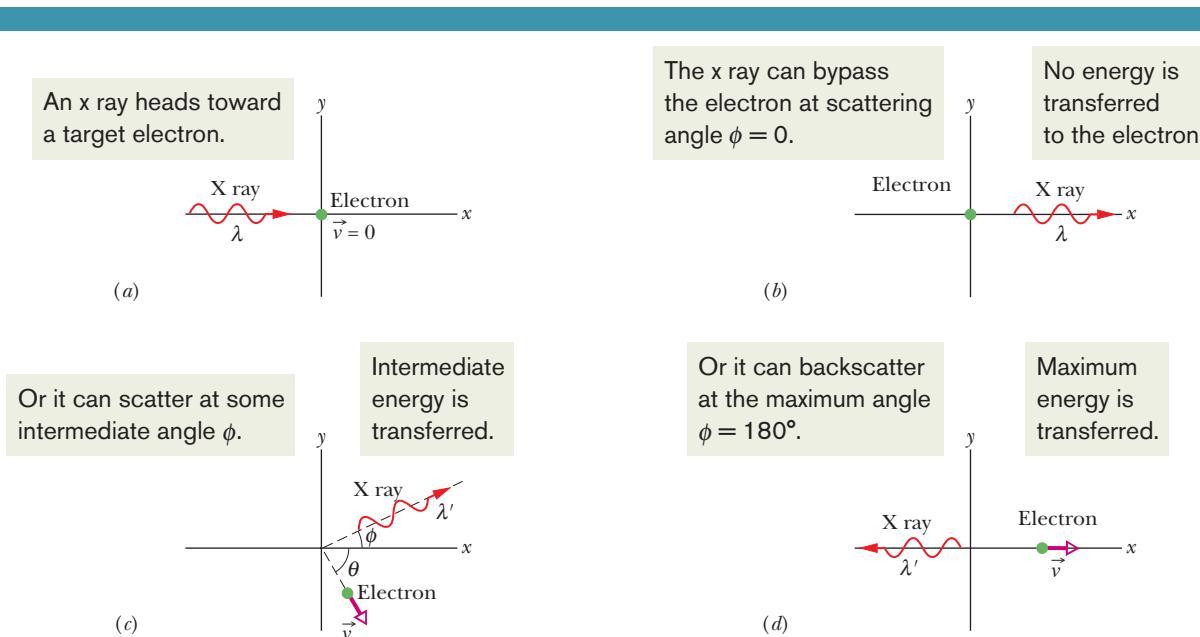


Fig. 38-5 (a) An x ray approaches a stationary electron. The x ray can (b) bypass the electron (forward scatter) with no energy or momentum transfer, (c) scatter at some intermediate angle with an intermediate energy and momentum transfer, or (d) backscatter with the maximum energy and momentum transfer.

in which hf is the energy of the incident x-ray photon, hf' is the energy of the scattered x-ray photon, and K is the kinetic energy of the recoiling electron. Because the electron may recoil with a speed comparable to that of light, we must use the relativistic expression of Eq. 37-52,

$$K = mc^2(\gamma - 1),$$

for the electron's kinetic energy. Here m is the electron's mass and γ is the Lorentz factor

$$\gamma = \frac{1}{\sqrt{1 - (v/c)^2}}.$$

Substituting for K in the conservation of energy equation yields

$$hf = hf' + mc^2(\gamma - 1).$$

Substituting c/λ for f and c/λ' for f' then leads to the new energy conservation equation

$$\frac{h}{\lambda} = \frac{h}{\lambda'} + mc(\gamma - 1). \quad (38-8)$$

Next we apply the law of conservation of momentum to the x-ray–electron collision of Fig. 38-5. From Eq. 38-7 ($p = h/\lambda$), the magnitude of the momentum of the incident photon is h/λ , and that of the scattered photon is h/λ' . From Eq. 37-41, the magnitude for the recoiling electron's momentum is $p = \gamma mv$. Because we have a two-dimensional situation, we write separate equations for the conservation of momentum along the x and y axes, obtaining

$$\frac{h}{\lambda} = \frac{h}{\lambda'} \cos \phi + \gamma mv \cos \theta \quad (x \text{ axis}) \quad (38-9)$$

and $0 = \frac{h}{\lambda'} \sin \phi - \gamma mv \sin \theta \quad (y \text{ axis}). \quad (38-10)$

We want to find $\Delta\lambda (= \lambda' - \lambda)$, the Compton shift of the scattered x rays. Of the five collision variables (λ , λ' , v , ϕ , and θ) that appear in Eqs. 38-8, 38-9, and 38-10, we choose to eliminate v and θ , which deal only with the recoiling electron. Carrying out the algebra (it is somewhat complicated) leads to

$$\Delta\lambda = \frac{h}{mc} (1 - \cos \phi) \quad (\text{Compton shift}). \quad (38-11)$$

Equation 38-11 agrees exactly with Compton's experimental results.

The quantity h/mc in Eq. 38-11 is a constant called the **Compton wavelength**. Its value depends on the mass m of the particle from which the x rays scatter. Here that particle is a loosely bound electron, and thus we would substitute the mass of an electron for m to evaluate the *Compton wavelength for Compton scattering from an electron*.

A Loose End

The peak at the incident wavelength $\lambda (= 71.1 \text{ pm})$ in Fig. 38-4 still needs to be explained. This peak arises not from interactions between x rays and the very loosely bound electrons in the target but from interactions between x rays and the electrons that are *tightly* bound to the carbon atoms making up the target. Effectively, each of these latter collisions occurs between an incident x ray and an entire carbon atom. If we substitute for m in Eq. 38-11 the mass of a carbon atom (which is about 22 000 times that of an electron), we see that $\Delta\lambda$ becomes about 22 000 times smaller than the Compton shift for an electron—too small to detect. Thus, the x rays scattered in these collisions have the same wavelength as the incident x rays.



CHECKPOINT 3

Compare Compton scattering for x rays ($\lambda \approx 20 \text{ pm}$) and visible light ($\lambda \approx 500 \text{ nm}$) at a particular angle of scattering. Which has the greater (a) Compton shift, (b) fractional wavelength shift, (c) fractional energy loss, and (d) energy imparted to the electron?

Sample Problem**Compton scattering of light by electrons**

X rays of wavelength $\lambda = 22 \text{ pm}$ (photon energy = 56 keV) are scattered from a carbon target, and the scattered rays are detected at 85° to the incident beam.

- (a) What is the Compton shift of the scattered rays?

KEY IDEA

The Compton shift is the wavelength change of the x rays due to scattering from loosely bound electrons in a target. Further, that shift depends on the angle at which the scattered x rays are detected, according to Eq. 38-11. The shift is zero for forward scattering at angle $\phi = 0^\circ$, and it is maximum for back scattering at angle $\phi = 180^\circ$. Here we have an intermediate situation at angle $\phi = 85^\circ$.

Calculation: Substituting 85° for that angle and $9.11 \times 10^{-31} \text{ kg}$ for the electron mass (because the scattering is from electrons) in Eq. 38-11 gives us

$$\begin{aligned}\Delta\lambda &= \frac{h}{mc}(1 - \cos\phi) \\ &= \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(1 - \cos 85^\circ)}{(9.11 \times 10^{-31} \text{ kg})(3.00 \times 10^8 \text{ m/s})} \\ &= 2.21 \times 10^{-12} \text{ m} \approx 2.2 \text{ pm}. \quad (\text{Answer})\end{aligned}$$

- (b) What percentage of the initial x-ray photon energy is transferred to an electron in such scattering?

KEY IDEA

We need to find the *fractional energy loss* (let us call it *frac*) for photons that scatter from the electrons:

$$\text{frac} = \frac{\text{energy loss}}{\text{initial energy}} = \frac{E - E'}{E}.$$

Calculations: From Eq. 38-2 ($E = hf$), we can substitute for the initial energy E and the detected energy E' of the x rays in terms of frequencies. Then, from Eq. 38-1 ($f = c/\lambda$), we can substitute for those frequencies in terms of the wavelengths. We find

$$\begin{aligned}\text{frac} &= \frac{hf - hf'}{hf} = \frac{c/\lambda - c/\lambda'}{c/\lambda} = \frac{\lambda' - \lambda}{\lambda'} \\ &= \frac{\Delta\lambda}{\lambda + \Delta\lambda}. \quad (38-12)\end{aligned}$$

Substitution of data yields

$$\text{frac} = \frac{2.21 \text{ pm}}{22 \text{ pm} + 2.21 \text{ pm}} = 0.091, \text{ or } 9.1\%. \quad (\text{Answer})$$

Although the Compton shift $\Delta\lambda$ is independent of the wavelength λ of the incident x rays (see Eq. 38-11), the *fractional photon energy loss* of the x rays does depend on λ , increasing as the wavelength of the incident radiation decreases, as indicated by Eq. 38-12.



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38-5 Light as a Probability Wave

A fundamental mystery in physics is how light can be a wave (which spreads out over a region) in classical physics but be emitted and absorbed as photons (which originate and vanish at points) in quantum physics. The double-slit experiment of Section 35-4 lies at the heart of this mystery. Let us discuss three versions of that experiment.

The Standard Version

Figure 38-6 is a sketch of the original experiment carried out by Thomas Young in 1801 (see also Fig. 35-8). Light shines on screen *B*, which contains two narrow parallel slits. The light waves emerging from the two slits spread out by diffraction and overlap on screen *C* where, by interference, they form a pattern of alternating intensity maxima and minima. In Section 35-4 we took the existence of these interference fringes as compelling evidence for the wave nature of light.

Let us place a tiny photon detector *D* at one point in the plane of screen *C*. Let the detector be a photoelectric device that clicks when it absorbs a photon. We would find that the detector produces a series of clicks, randomly spaced

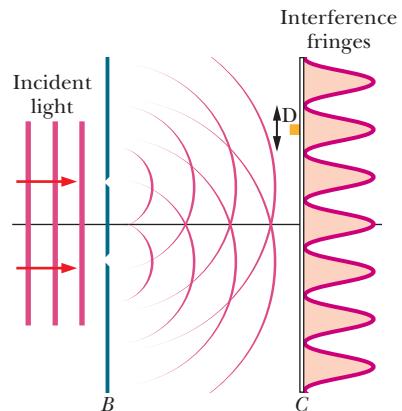


Fig. 38-6 Light is directed onto screen *B*, which contains two parallel slits. Light emerging from these slits spreads out by diffraction. The two diffracted waves overlap at screen *C* and form a pattern of interference fringes. A small photon detector *D* in the plane of screen *C* generates a sharp click for each photon that it absorbs.

in time, each click signaling the transfer of energy from the light wave to the screen via a photon absorption. If we moved the detector very slowly up or down as indicated by the black arrow in Fig. 38-6, we would find that the click rate increases and decreases, passing through alternate maxima and minima that correspond exactly to the maxima and minima of the interference fringes.

The point of this thought experiment is as follows. We cannot predict when a photon will be detected at any particular point on screen *C*; photons are detected at individual points at random times. We can, however, predict that the relative *probability* that a single photon will be detected at a particular point in a specified time interval is proportional to the light intensity at that point.

We know from Eq. 33-26 ($I = E_{\text{rms}}^2/c\mu_0$) in Section 33-5 that the intensity *I* of a light wave at any point is proportional to the square of E_m , the amplitude of the oscillating electric field vector of the wave at that point. Thus,



The probability (per unit time interval) that a photon will be detected in any small volume centered on a given point in a light wave is proportional to the square of the amplitude of the wave's electric field vector at that point.

We now have a probabilistic description of a light wave, hence another way to view light. It is not only an electromagnetic wave but also a **probability wave**. That is, to every point in a light wave we can attach a numerical probability (per unit time interval) that a photon can be detected in any small volume centered on that point.

The Single-Photon Version

A single-photon version of the double-slit experiment was first carried out by G. I. Taylor in 1909 and has been repeated many times since. It differs from the standard version in that the light source in the Taylor experiment is so extremely feeble that it emits only one photon at a time, at random intervals. Astonishingly, interference fringes still build up on screen *C* if the experiment runs long enough (several months for Taylor's early experiment).

What explanation can we offer for the result of this single-photon double-slit experiment? Before we can even consider the result, we are compelled to ask questions like these: If the photons move through the apparatus one at a time, through which of the two slits in screen *B* does a given photon pass? How does a given photon even "know" that there is another slit present so that interference is a possibility? Can a single photon somehow pass through both slits and interfere with itself?

Bear in mind that the only thing we can know about photons is when light interacts with matter—we have no way of detecting them without an interaction with matter, such as with a detector or a screen. Thus, in the experiment of Fig. 38-6, all we can know is that photons originate at the light source and vanish at the screen. Between source and screen, we cannot know what the photon is or does. However, because an interference pattern eventually builds up on the screen, we can speculate that each photon travels from source to screen *as a wave* that fills up the space between source and screen and then vanishes in a photon absorption at some point on the screen, with a transfer of energy and momentum to the screen at that point.

We *cannot* predict where this transfer will occur (where a photon will be detected) for any given photon originating at the source. However, we *can* predict the probability that a transfer will occur at any given point on the screen. Transfers will tend to occur (and thus photons will tend to be absorbed) in the regions of the bright fringes in the interference pattern that builds up on the screen. Transfers will tend *not* to occur (and thus photons will tend *not* to be absorbed) in the regions of the dark fringes in the built-up pattern. Thus, we can say that the wave traveling from the source to the screen is a *probability wave*, which produces a pattern of "probability fringes" on the screen.

The Single-Photon, Wide-Angle Version

In the past, physicists tried to explain the single-photon double-slit experiment in terms of small packets of classical light waves that are individually sent toward the slits. They would define these small packets as photons. However, modern experiments invalidate this explanation and definition. Figure 38-7 shows the arrangement of one of these experiments, reported in 1992 by Ming Lai and Jean-Claude Diels of the University of New Mexico. Source S contains molecules that emit photons at well separated times. Mirrors M_1 and M_2 are positioned to reflect light that the source emits along two distinct paths, 1 and 2, that are separated by an angle θ , which is close to 180° . This arrangement differs from the standard two-slit experiment, in which the angle between the paths of the light reaching two slits is very small.

After reflection from mirrors M_1 and M_2 , the light waves traveling along paths 1 and 2 meet at beam splitter B. (A beam splitter is an optical device that transmits half the light incident upon it and reflects the other half.) On the right side of the beam splitter in Fig. 38-7, the light wave traveling along path 2 and reflected by B combines with the light wave traveling along path 1 and transmitted by B. These two waves then interfere with each other as they arrive at detector D (a *photomultiplier tube* that can detect individual photons).

The output of the detector is a randomly spaced series of electronic pulses, one for each detected photon. In the experiment, the beam splitter is moved slowly in a horizontal direction (in the reported experiment, a distance of only about $50\ \mu\text{m}$ maximum), and the detector output is recorded on a chart recorder. Moving the beam splitter changes the lengths of paths 1 and 2, producing a phase shift between the light waves arriving at detector D. Interference maxima and minima appear in the detector's output signal.

This experiment is difficult to understand in traditional terms. For example, when a molecule in the source emits a single photon, does that photon travel along path 1 or path 2 in Fig. 38-7 (or along any other path)? Or can it move in both directions at once? To answer, we assume that when a molecule emits a photon, a probability wave radiates in all directions from it. The experiment samples this wave in two of those directions, chosen to be nearly opposite each other.

We see that we can interpret all three versions of the double-slit experiment if we assume that (1) light is generated in the source as photons, (2) light is absorbed in the detector as photons, and (3) light travels between source and detector as a probability wave.

38-6 Electrons and Matter Waves

In 1924, French physicist Louis de Broglie made the following appeal to symmetry: A beam of light is a wave, but it transfers energy and momentum to matter only at points, via photons. Why can't a beam of particles have the same properties? That is, why can't we think of a moving electron—or any other particle—as a **matter wave** that transfers energy and momentum to other matter at points?

In particular, de Broglie suggested that Eq. 38-7 ($p = h/\lambda$) might apply not only to photons but also to electrons. We used that equation in Section 38-4 to assign a momentum p to a photon of light with wavelength λ . We now use it, in the form

$$\lambda = \frac{h}{p} \quad (\text{de Broglie wavelength}), \quad (38-13)$$

to assign a wavelength λ to a particle with momentum of magnitude p . The wavelength calculated from Eq. 38-13 is called the **de Broglie wavelength** of the moving particle. De Broglie's prediction of the existence of matter waves was first verified experimentally in 1927, by C. J. Davisson and L. H. Germer of the

A single photon can take widely different paths and still interfere with itself.

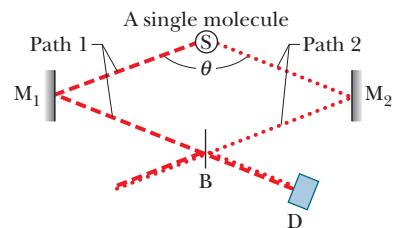


Fig. 38-7 The light from a single photon emission in source S travels over two widely separated paths and interferes with itself at detector D after being recombined by beam splitter B. (After Ming Lai and Jean-Claude Diels, *Journal of the Optical Society of America B*, **9**, 2290–2294, December 1992.)

Bell Telephone Laboratories and by George P. Thomson of the University of Aberdeen in Scotland.

Figure 38-8 shows photographic proof of matter waves in a more recent experiment. In the experiment, an interference pattern was built up when electrons were sent, *one by one*, through a double-slit apparatus. The apparatus was like the ones we have previously used to demonstrate optical interference, except that the viewing screen was similar to an old-fashioned television screen. When an electron hit the screen, it caused a flash of light whose position was recorded.

The first several electrons (top two photos) revealed nothing interesting and seemingly hit the screen at random points. However, after many thousands of electrons were sent through the apparatus, a pattern appeared on the screen, revealing fringes where many electrons had hit the screen and fringes where few had hit the screen. The pattern is exactly what we would expect for wave interference. Thus, *each* electron passed through the apparatus as a matter wave—the portion of the matter wave that traveled through one slit interfered with the portion that traveled through the other slit. That interference then determined the probability that the electron would materialize at a given point on the screen, hitting the screen there. Many electrons materialized in regions corresponding to bright fringes in optical interference, and few electrons materialized in regions corresponding to dark fringes.

Similar interference has been demonstrated with protons, neutrons, and various atoms. In 1994, it was demonstrated with iodine molecules I_2 , which are not only 500 000 times more massive than electrons but far more complex. In 1999, it was demonstrated with the even more complex *fullerenes* (or *buckyballs*) C_{60} .

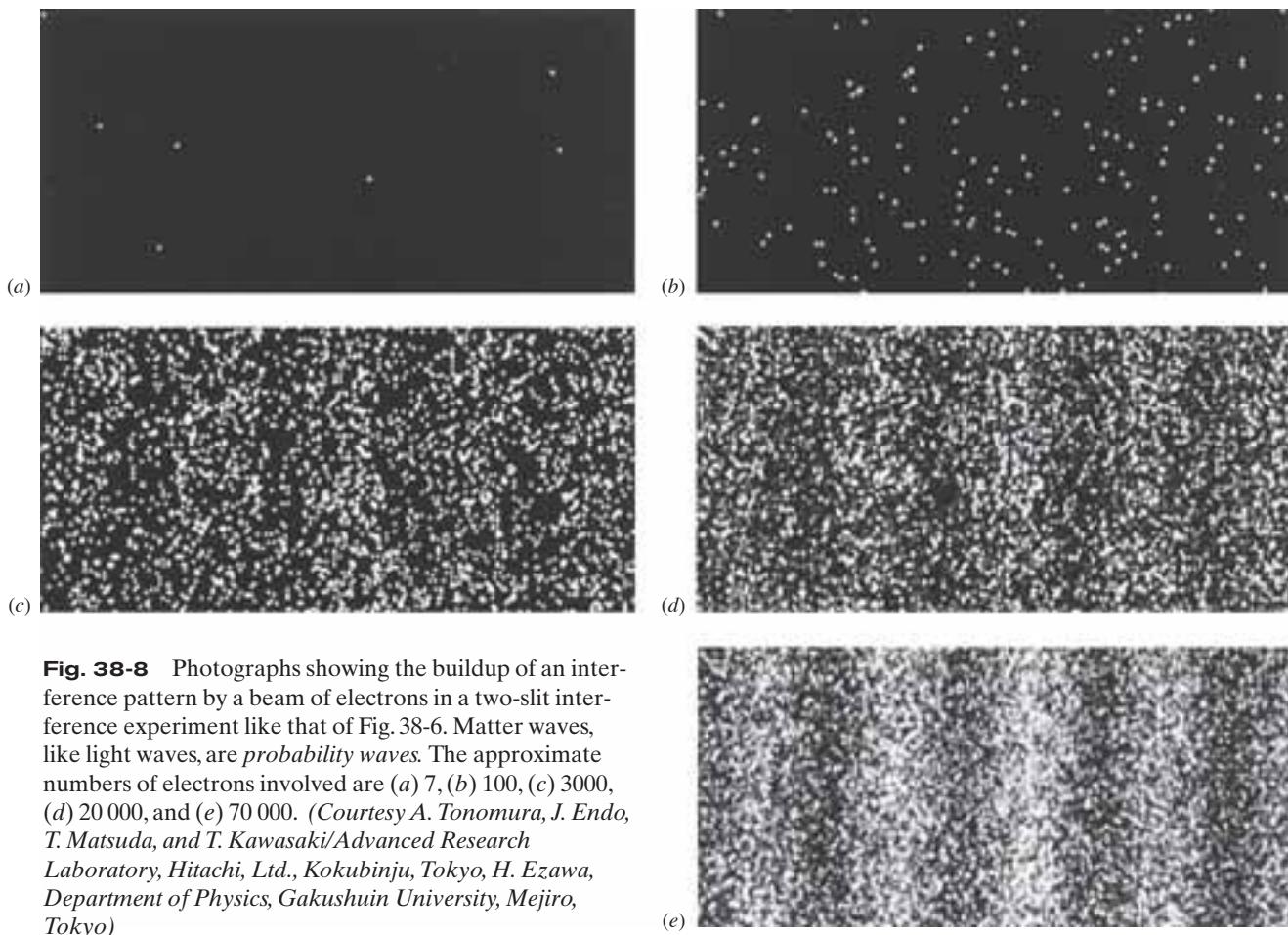
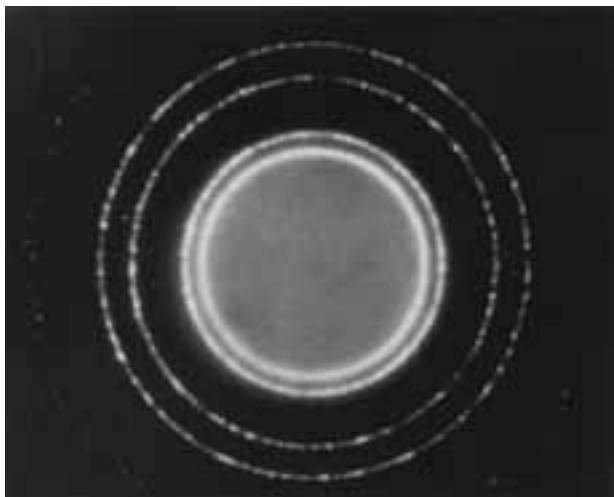
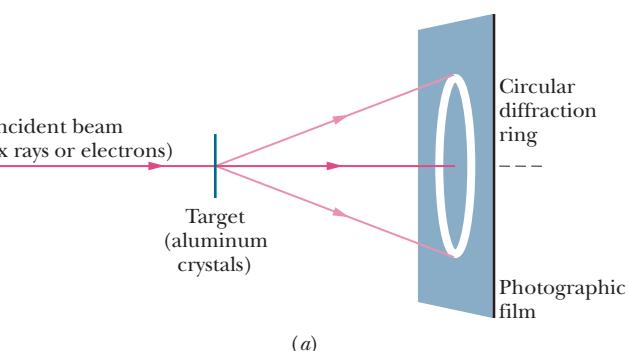


Fig. 38-8 Photographs showing the buildup of an interference pattern by a beam of electrons in a two-slit interference experiment like that of Fig. 38-6. Matter waves, like light waves, are *probability waves*. The approximate numbers of electrons involved are (a) 7, (b) 100, (c) 3000, (d) 20 000, and (e) 70 000. (*Courtesy A. Tonomura, J. Endo, T. Matsuda, and T. Kawasaki/Advanced Research Laboratory, Hitachi, Ltd., Kokubinju, Tokyo, H. Ezawa, Department of Physics, Gakushuin University, Mejiro, Tokyo*)

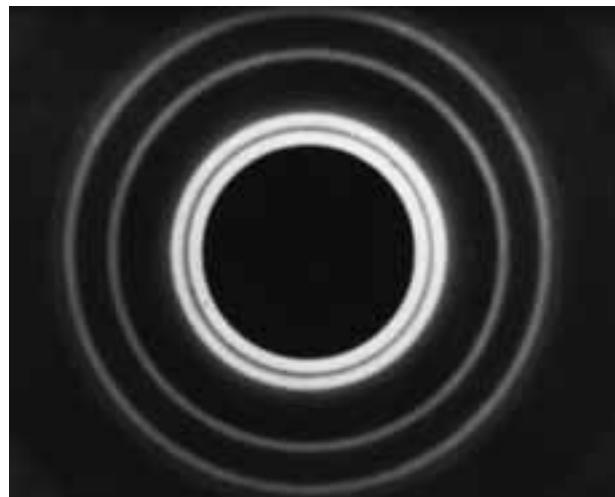
Fig. 38-9 (a) An experimental arrangement used to demonstrate, by diffraction techniques, the wave-like character of the incident beam. Photographs of the diffraction patterns when the incident beam is (b) an x-ray beam (light wave) and (c) an electron beam (matter wave). Note that the two patterns are geometrically identical to each other. (Photo (b) Cameca, Inc. Photo (c) from PSSC film "Matter Waves," courtesy Education Development Center, Newton, Massachusetts)



(b)



(a)



(c)

and C_{70} . (Fullerenes are molecules of carbon atoms that are arranged in a structure resembling a soccer ball, 60 carbon atoms in C_{60} and 70 carbon atoms in C_{70} .) Apparently, such small objects as electrons, protons, atoms, and molecules travel as matter waves. However, as we consider larger and more complex objects, there must come a point at which we are no longer justified in considering the wave nature of an object. At that point, we are back in our familiar nonquantum world, with the physics of earlier chapters of this book. In short, an electron is a matter wave and can undergo interference with itself, but a cat is not a matter wave and cannot undergo interference with itself (which must be a relief to cats).

The wave nature of particles and atoms is now taken for granted in many scientific and engineering fields. For example, electron diffraction and neutron diffraction are used to study the atomic structures of solids and liquids, and electron diffraction is used to study the atomic features of surfaces on solids.

Figure 38-9a shows an arrangement that can be used to demonstrate the scattering of either x rays or electrons by crystals. A beam of one or the other is directed onto a target consisting of a layer of tiny aluminum crystals. The x rays have a certain wavelength λ . The electrons are given enough energy so that their de Broglie wavelength is the same wavelength λ . The scatter of x rays or electrons by the crystals produces a circular interference pattern on a photographic film. Figure 38-9b shows the pattern for the scatter of x rays, and Fig. 38-9c shows the pattern for the scatter of electrons. The patterns are the same—both x rays and electrons are waves.

Waves and Particles

Figures 38-8 and 38-9 are convincing evidence of the *wave* nature of matter, but we have countless experiments that suggest its *particle* nature. Figure 38-10, for

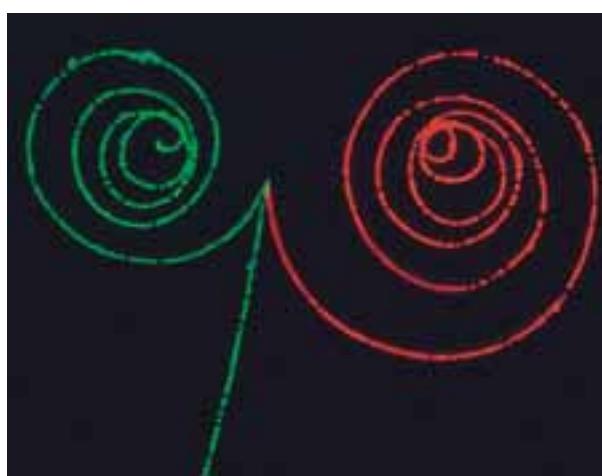


Fig. 38-10 A bubble-chamber image showing where two electrons (paths color coded green) and one positron (red) moved after a gamma ray entered the chamber. (Lawrence Berkeley Laboratory/Science Photo Library/Photo Researchers)

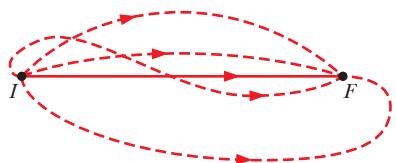


Fig. 38-11 A few of the many paths that connect two particle detection points I and F . Only matter waves that follow paths close to the straight line between these points interfere constructively. For all other paths, the waves following any pair of neighboring paths interfere destructively. Thus, a matter wave leaves a straight track.

example, shows the tracks of particles (rather than waves) revealed in a bubble chamber. When a charged particle passes through the liquid hydrogen that fills such a chamber, the particle causes the liquid to vaporize along the particle's path. A series of bubbles thus marks the path, which is usually curved due to a magnetic field set up perpendicular to the plane of the chamber.

In Fig. 38-10, a gamma ray left no track when it entered at the top because the ray is electrically neutral and thus caused no vapor bubbles as it passed through the liquid hydrogen. However, it collided with one of the hydrogen atoms, kicking an electron out of that atom; the curved path taken by the electron to the bottom of the photograph has been color coded green. Simultaneous with the collision, the gamma ray transformed into an electron and a positron in a pair production event (see Eq. 21-15). Those two particles then moved in tight spirals (color coded green for the electron and red for the positron) as they gradually lost energy in repeated collisions with hydrogen atoms. Surely these tracks are evidence of the particle nature of the electron and positron, but is there any evidence of waves in Fig. 38-10?

To simplify the situation, let us turn off the magnetic field so that the strings of bubbles will be straight. We can view each bubble as a detection point for the electron. Matter waves traveling between detection points such as I and F in Fig. 38-11 will explore all possible paths, a few of which are shown.

In general, for every path connecting I and F (except the straight-line path), there will be a neighboring path such that matter waves following the two paths cancel each other by interference. This is not true, however, for the straight-line path joining I and F ; in this case, matter waves traversing all neighboring paths reinforce the wave following the direct path. You can think of the bubbles that form the track as a series of detection points at which the matter wave undergoes constructive interference.



CHECKPOINT 4

For an electron and a proton that have the same (a) kinetic energy, (b) momentum, or (c) speed, which particle has the shorter de Broglie wavelength?

Sample Problem

de Broglie wavelength of an electron

What is the de Broglie wavelength of an electron with a kinetic energy of 120 eV?

KEY IDEAS

(1) We can find the electron's de Broglie wavelength λ from Eq. 38-13 ($\lambda = h/p$) if we first find the magnitude of its momentum p . (2) We find p from the given kinetic energy K of the electron. That kinetic energy is much less than the rest energy of an electron (0.511 MeV, from Table 37-3). Thus, we can get by with the classical approximations for momentum p ($= mv$) and kinetic energy K ($= \frac{1}{2}mv^2$).

Calculations: We are given the value of the kinetic energy. So, in order to use the de Broglie relation, we first solve the kinetic energy equation for v and then substitute into the

momentum equation, finding

$$\begin{aligned} p &= \sqrt{2mK} \\ &= \sqrt{(2)(9.11 \times 10^{-31} \text{ kg})(120 \text{ eV})(1.60 \times 10^{-19} \text{ J/eV})} \\ &= 5.91 \times 10^{-24} \text{ kg}\cdot\text{m/s}. \end{aligned}$$

From Eq. 38-13 then

$$\begin{aligned} \lambda &= \frac{h}{p} \\ &= \frac{6.63 \times 10^{-34} \text{ J}\cdot\text{s}}{5.91 \times 10^{-24} \text{ kg}\cdot\text{m/s}} \\ &= 1.12 \times 10^{-10} \text{ m} = 112 \text{ pm}. \quad (\text{Answer}) \end{aligned}$$

This wavelength associated with the electron is about the size of a typical atom. If we increase the electron's kinetic energy, the wavelength becomes even smaller.



Additional examples, video, and practice available at WileyPLUS

38-7 Schrödinger's Equation

A simple traveling wave of any kind, be it a wave on a string, a sound wave, or a light wave, is described in terms of some quantity that varies in a wave-like fashion. For light waves, for example, this quantity is $\vec{E}(x, y, z, t)$, the electric field component of the wave. Its observed value at any point depends on the location of that point and on the time at which the observation is made.

What varying quantity should we use to describe a matter wave? We should expect this quantity, which we call the **wave function** $\Psi(x, y, z, t)$, to be more complicated than the corresponding quantity for a light wave because a matter wave, in addition to energy and momentum, transports mass and (often) electric charge. It turns out that Ψ , the uppercase Greek letter psi, usually represents a function that is complex in the mathematical sense; that is, we can always write its values in the form $a + ib$, in which a and b are real numbers and $i^2 = -1$.

In all the situations you will meet here, the space and time variables can be grouped separately and Ψ can be written in the form

$$\Psi(x, y, z, t) = \psi(x, y, z) e^{-i\omega t}, \quad (38-14)$$

where $\omega (= 2\pi f)$ is the angular frequency of the matter wave. Note that ψ , the lowercase Greek letter psi, represents only the space-dependent part of the complete, time-dependent wave function Ψ . We shall focus on ψ . Two questions arise: What is meant by the wave function? How do we find it?

What does the wave function mean? It has to do with the fact that a matter wave, like a light wave, is a probability wave. Suppose that a matter wave reaches a particle detector that is small; then the probability that a particle will be detected in a specified time interval is proportional to $|\psi|^2$, where $|\psi|$ is the absolute value of the wave function at the location of the detector. Although ψ is usually a complex quantity, $|\psi|^2$ is always both real and positive. It is, then, $|\psi|^2$, which we call the **probability density**, and not ψ , that has *physical* meaning. Speaking loosely, the meaning is this:



The probability (per unit time) of detecting a particle in a small volume centered on a given point in a matter wave is proportional to the value of $|\psi|^2$ at that point.

Because ψ is usually a complex quantity, we find the square of its absolute value by multiplying ψ by ψ^* , the *complex conjugate* of ψ . (To find ψ^* we replace the imaginary number i in ψ with $-i$, wherever it occurs.)

How do we find the wave function? Sound waves and waves on strings are described by the equations of Newtonian mechanics. Light waves are described by Maxwell's equations. Matter waves are described by **Schrödinger's equation**, advanced in 1926 by Austrian physicist Erwin Schrödinger.

Many of the situations that we shall discuss involve a particle traveling in the x direction through a region in which forces acting on the particle cause it to have a potential energy $U(x)$. In this special case, Schrödinger's equation reduces to

$$\frac{d^2\psi}{dx^2} + \frac{8\pi^2m}{h^2} [E - U(x)]\psi = 0 \quad (\text{Schrödinger's equation, one-dimensional motion}), \quad (38-15)$$

in which E is the total mechanical energy of the moving particle. (We do *not* consider mass energy in this nonrelativistic equation.) We cannot derive Schrödinger's equation from more basic principles; it *is* the basic principle.

If $U(x)$ in Eq. 38-15 is zero, that equation describes a **free particle**—that is, a moving particle on which no net force acts. The particle's total energy in this case is all kinetic, and thus E in Eq. 38-15 is $\frac{1}{2}mv^2$. That equation then becomes

$$\frac{d^2\psi}{dx^2} + \frac{8\pi^2m}{h^2} \left(\frac{mv^2}{2} \right) \psi = 0,$$

which we can recast as

$$\frac{d^2\psi}{dx^2} + \left(2\pi \frac{p}{h}\right)^2 \psi = 0.$$

To obtain this equation, we replaced mv with the momentum p and regrouped terms.

From Eq. 38-13 ($\lambda = h/p$) we recognize p/h in the equation above as $1/\lambda$, where λ is the de Broglie wavelength of the moving particle. We further recognize $2\pi/\lambda$ as the *angular wave number* k , which we defined in Eq. 16-5. With these substitutions, the equation above becomes

$$\frac{d^2\psi}{dx^2} + k^2\psi = 0 \quad (\text{Schrödinger's equation, free particle}). \quad (38-16)$$

The most general solution of Eq. 38-16 is

$$\psi(x) = Ae^{ikx} + Be^{-ikx}, \quad (38-17)$$

in which A and B are arbitrary constants. You can show that this equation is indeed a solution of Eq. 38-16 by substituting $\psi(x)$ and its second derivative into that equation and noting that an identity results.

If we combine Eqs. 38-14 and 38-17, we find, for the time-dependent wave function Ψ of a free particle traveling in the x direction,

$$\begin{aligned} \Psi(x, t) &= \psi(x)e^{-i\omega t} = (Ae^{ikx} + Be^{-ikx})e^{-i\omega t} \\ &= Ae^{i(kx - \omega t)} + Be^{-i(kx + \omega t)}. \end{aligned} \quad (38-18)$$

Finding the Probability Density $|\psi|^2$

In Section 16-5 we saw that *any function* F of the form $F(kx \pm \omega t)$ represents a traveling wave. This applies to exponential functions like those in Eq. 38-18 as well as to the sinusoidal functions we have used to describe waves on strings. For a general angle θ , these two representations of functions are related by

$$e^{i\theta} = \cos \theta + i \sin \theta \quad \text{and} \quad e^{-i\theta} = \cos \theta - i \sin \theta.$$

The first term on the right in Eq. 38-18 thus represents a wave traveling in the positive direction of x and the second term represents a wave traveling in the negative direction of x . However, we have assumed that the free particle we are considering travels only in the positive direction of x . To reduce the general solution (Eq. 38-18) to our case of interest, we choose the arbitrary constant B in Eqs. 38-18 and 38-17 to be zero. At the same time, we relabel the constant A as ψ_0 . Equation 38-17 then becomes

$$\psi(x) = \psi_0 e^{ikx}. \quad (38-19)$$

To calculate the probability density, we take the square of the absolute value:

$$|\psi|^2 = |\psi_0 e^{ikx}|^2 = (\psi_0^2)|e^{ikx}|^2.$$

Now, because

$$|e^{ikx}|^2 = (e^{ikx})(e^{ikx})^* = e^{ikx} e^{-ikx} = e^{ikx - ikx} = e^0 = 1,$$

we get

$$|\psi|^2 = (\psi_0^2)(1)^2 = \psi_0^2 \quad (\text{a constant}).$$

Figure 38-12 is a plot of the probability density $|\psi|^2$ versus x for a free particle—a straight line parallel to the x axis from $-\infty$ to $+\infty$. We see that the probability density $|\psi|^2$ is the same for all values of x , which means that the particle has equal probabilities of being *anywhere* along the x axis. There is no distinguishing feature by which we can predict a most likely position for the particle. That is, all positions are equally likely. We'll see what this means in the next section.

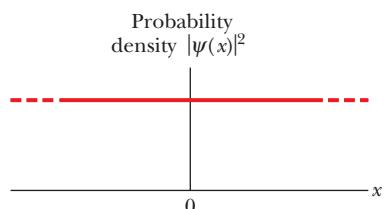


Fig. 38-12 A plot of the probability density $|\psi|^2$ for a free particle moving in the positive x direction. Since $|\psi|^2$ has the same constant value for all values of x , the particle has the same probability of detection at all points along its path.

38-8 Heisenberg's Uncertainty Principle

Our inability to predict the position of a free particle, as indicated by Fig. 38-12, is our first example of **Heisenberg's uncertainty principle**, proposed in 1927 by German physicist Werner Heisenberg. It states that measured values cannot be assigned to the position \vec{r} and the momentum \vec{p} of a particle simultaneously with unlimited precision.

In terms of $\hbar = h/2\pi$ (called “h-bar”), the principle tells us

$$\begin{aligned}\Delta x \cdot \Delta p_x &\geq \hbar \\ \Delta y \cdot \Delta p_y &\geq \hbar \quad (\text{Heisenberg's uncertainty principle).} \\ \Delta z \cdot \Delta p_z &\geq \hbar\end{aligned}\quad (38-20)$$

Here Δx and Δp_x represent the intrinsic uncertainties in the measurements of the x components of \vec{r} and \vec{p} , with parallel meanings for the y and z terms. Even with the best measuring instruments, each product of a position uncertainty and a momentum uncertainty in Eq. 38-20 will be greater than \hbar , *never less*.

The particle whose probability density is plotted in Fig. 38-12 is a free particle; that is, no force acts on it, and so its momentum \vec{p} must be constant. We implied—without making a point of it—that we can determine \vec{p} with absolute precision; in other words, we assumed that $\Delta p_x = \Delta p_y = \Delta p_z = 0$ in Eq. 38-20. That assumption then requires $\Delta x \rightarrow \infty$, $\Delta y \rightarrow \infty$, and $\Delta z \rightarrow \infty$. With such infinitely great uncertainties, the position of the particle is completely unspecified.

Do not think that the particle *really has* a sharply defined position that is, for some reason, hidden from us. If its momentum can be specified with absolute precision, the words “position of the particle” simply lose all meaning. The particle in Fig. 38-12 can be found *with equal probability* anywhere along the x axis.

Sample Problem

Uncertainty principle: position and momentum

Assume that an electron is moving along an x axis and that you measure its speed to be $2.05 \times 10^6 \text{ m/s}$, which can be known with a precision of 0.50%. What is the minimum uncertainty (as allowed by the uncertainty principle in quantum theory) with which you can simultaneously measure the position of the electron along the x axis?

KEY IDEA

The minimum uncertainty allowed by quantum theory is given by Heisenberg's uncertainty principle in Eq. 38-20. We need only consider components along the x axis because we have motion only along that axis and want the uncertainty Δx in location along that axis. Since we want the minimum allowed uncertainty, we use the equality instead of the inequality in the x -axis part of Eq. 38-20, writing $\Delta x \cdot \Delta p_x = \hbar$.

Calculations: To evaluate the uncertainty Δp_x in the momentum, we must first evaluate the momentum component p_x . Because the electron's speed v_x is much less than the speed of light c , we can evaluate p_x with the classical expression for momentum instead of using a relativistic expres-

sion. We find

$$\begin{aligned}p_x &= mv_x = (9.11 \times 10^{-31} \text{ kg})(2.05 \times 10^6 \text{ m/s}) \\ &= 1.87 \times 10^{-24} \text{ kg} \cdot \text{m/s}.\end{aligned}$$

The uncertainty in the speed is given as 0.50% of the measured speed. Because p_x depends directly on speed, the uncertainty Δp_x in the momentum must be 0.50% of the momentum:

$$\begin{aligned}\Delta p_x &= (0.0050)p_x \\ &= (0.0050)(1.87 \times 10^{-24} \text{ kg} \cdot \text{m/s}) \\ &= 9.35 \times 10^{-27} \text{ kg} \cdot \text{m/s}.\end{aligned}$$

Then the uncertainty principle gives us

$$\begin{aligned}\Delta x &= \frac{\hbar}{\Delta p_x} = \frac{(6.63 \times 10^{-34} \text{ J} \cdot \text{s})/2\pi}{9.35 \times 10^{-27} \text{ kg} \cdot \text{m/s}} \\ &= 1.13 \times 10^{-8} \text{ m} \approx 11 \text{ nm}, \quad (\text{Answer})\end{aligned}$$

which is about 100 atomic diameters. Given your measurement of the electron's speed, it makes no sense to try to pin down the electron's position to any greater precision.



Additional examples, video, and practice available at WileyPLUS

38-9 Barrier Tunneling

Suppose you slide a puck over frictionless ice toward an ice-covered hill (Fig. 38-13). As the puck climbs the hill, kinetic energy K is transformed into gravitational potential energy U . If the puck reaches the top, its potential energy is U_b . Thus, the puck can pass over the top only if its initial mechanical energy $E > U_b$. Otherwise, the puck eventually stops its climb up the left side of the hill and slides back to the left. For instance, if $U_b = 20 \text{ J}$ and $E = 10 \text{ J}$, you cannot expect the puck to pass over the hill. We say that the hill acts as a **potential energy barrier** (or, for short, a **potential barrier**) and that, in this case, the barrier has a **height** of $U_b = 20 \text{ J}$.

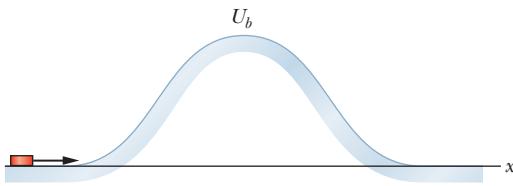


Fig. 38-13 A puck slides over frictionless ice toward a hill. The puck's gravitational potential at the top of the hill will be U_b .

Figure 38-14 shows a potential barrier for a nonrelativistic electron traveling along an idealized wire of negligible thickness. The electron, with mechanical energy E , approaches a region (the barrier) in which the electric potential V_b is negative. Because it is negatively charged, the electron will have a positive potential energy $U_b (= qV_b)$ in that region (Fig. 38-15). If $E > U_b$, we expect the electron to pass through the barrier region and come out to the right of $x = L$ in Fig. 38-14. Nothing surprising there. If $E < U_b$, we expect the electron to be unable to pass through the barrier region. Instead, it should end up traveling leftward, much as the puck would slide back down the hill in Fig. 38-13 if the puck has $E < U_b$.

However, something astounding can happen to the electron when $E < U_b$. Because it is a matter wave, the electron has a finite probability of leaking (or, better, *tunneling*) through the barrier and materializing on the other side, moving rightward with energy E as though nothing (strange or otherwise) had happened in the region of $0 \leq x \leq L$.

The wave function $\psi(x)$ describing the electron can be found by solving Schrödinger's equation (Eq. 38-15) separately for the three regions in Fig. 38-14: (1) to the left of the barrier, (2) within the barrier, and (3) to the right of the barrier. The arbitrary constants that appear in the solutions can then be chosen so that the values of $\psi(x)$ and its derivative with respect to x join smoothly (no jumps, no kinks) at $x = 0$ and at $x = L$. Squaring the absolute value of $\psi(x)$ then yields the probability density.

Figure 38-16 shows a plot of the result. The oscillating curve to the left of the barrier (for $x < 0$) is a combination of the incident matter wave and the reflected matter wave (which has a smaller amplitude than the incident wave). The oscilla-

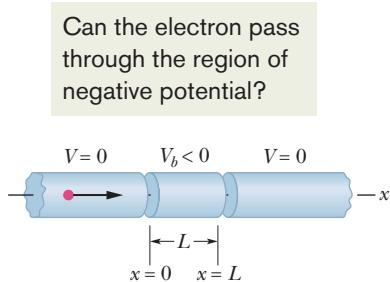


Fig. 38-14 The elements of an idealized thin wire in which an electron (the dot) approaches a negative electric potential V_b in the region $x = 0$ to $x = L$.

Classically, the electron lacks the energy to pass through the barrier region.

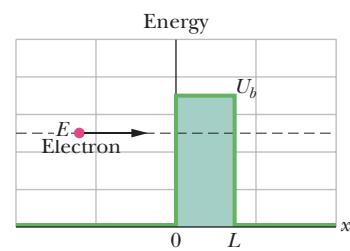


Fig. 38-15 An energy diagram containing two plots for the situation of Fig. 38-13: (1) The electron's mechanical energy E is plotted when the electron is at any coordinate $x < 0$. (2) The electron's electric potential energy U is plotted as a function of the electron's position x , assuming that the electron can reach any value of x . The nonzero part of the plot (the potential barrier) has height U_b and thickness L .

38-9 BARRIER TUNNELING

1075

tions occur because these two waves, traveling in opposite directions, interfere with each other, setting up a standing wave pattern.

Within the barrier (for $0 < x < L$) the probability density decreases exponentially with x . However, if L is small, the probability density is not quite zero at $x = L$.

To the right of the barrier (for $x > L$), the probability density plot describes a transmitted (through the barrier) wave with low but constant amplitude. Thus, the electron can be detected in this region but with a relatively small probability. (Compare this part of the figure with Fig. 38-12 for a free particle.)

We can assign a **transmission coefficient** T to the incident matter wave and the barrier. This coefficient gives the probability with which an approaching electron will be transmitted through the barrier—that is, that tunneling will occur. As an example, if $T = 0.020$, then of every 1000 electrons fired at the barrier, 20 (on average) will tunnel through it and 980 will be reflected. The transmission coefficient T is approximately

$$T \approx e^{-2bL}, \quad (38-21)$$

in which

$$b = \sqrt{\frac{8\pi^2 m(U_b - E)}{h^2}}, \quad (38-22)$$

and e is the exponential function. Because of the exponential form of Eq. 38-21, the value of T is very sensitive to the three variables on which it depends: particle mass m , barrier thickness L , and energy difference $U_b - E$. (Because we do not include relativistic effects here, E does not include mass energy.)

Barrier tunneling finds many applications in technology, including the tunnel diode, in which a flow of electrons produced by tunneling can be rapidly turned on or off by controlling the barrier height. The 1973 Nobel Prize in physics was shared by three “tunnelers,” Leo Esaki (for tunneling in semiconductors), Ivar Giaever (for tunneling in superconductors), and Brian Josephson (for the Josephson junction, a rapid quantum switching device based on tunneling). The 1986 Nobel Prize was awarded to Gerd Binnig and Heinrich Rohrer for development of the scanning tunneling microscope.



CHECKPOINT 5

Is the wavelength of the transmitted wave in Fig. 38-16 larger than, smaller than, or the same as that of the incident wave?

The Scanning Tunneling Microscope (STM)

The size of details that can be seen in an optical microscope is limited by the wavelength of the light the microscope uses (about 300 nm for ultraviolet light). The size of details that can be seen in the image that opens this chapter is far smaller and thus requires much smaller wavelengths. The waves used are electron matter waves, but they do not scatter from the surface being examined the way waves do in an optical microscope. Instead, the images we see are created by electrons tunneling through potential barriers at the tip of a *scanning tunneling microscope* (STM).

Figure 38-17 shows the heart of the scanning tunneling microscope. A fine metallic tip, mounted at the intersection of three mutually perpendicular quartz rods, is placed close to the surface to be examined. A small potential difference, perhaps only 10 mV, is applied between tip and surface.

Crystalline quartz has an interesting property called *piezoelectricity*: When an electric potential difference is applied across a sample of crystalline quartz, the dimensions of the sample change slightly. This property is used to change the length of each of the three rods in Fig. 38-17, smoothly and by tiny amounts, so that the tip can be scanned back and forth over the surface (in the x and y directions) and also lowered or raised with respect to the surface (in the z direction).

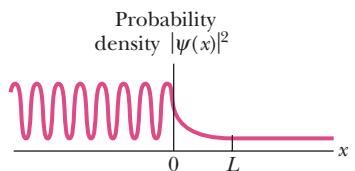


Fig. 38-16 A plot of the probability density $|\psi|^2$ of the electron matter wave for the situation of Fig. 38-15. The value of $|\psi|^2$ is nonzero to the right of the potential barrier.

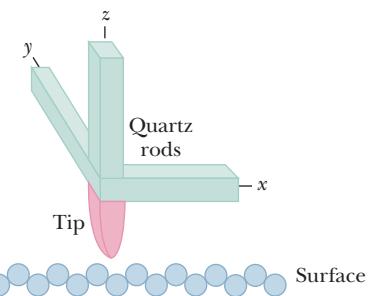


Fig. 38-17 The essence of a scanning tunneling microscope (STM). Three quartz rods are used to scan a sharply pointed conducting tip across the surface of interest and to maintain a constant separation between tip and surface. The tip thus moves up and down to match the contours of the surface, and a record of its movement provides information for a computer to create an image of the surface.

The space between the surface and the tip forms a potential energy barrier, much like that plotted in Fig. 38-15. If the tip is close enough to the surface, electrons from the sample can tunnel through this barrier from the surface to the tip, forming a tunneling current.

In operation, an electronic feedback arrangement adjusts the vertical position of the tip to keep the tunneling current constant as the tip is scanned over the surface. This means that the tip–surface separation also remains constant during the scan. The output of the device is a video display of the varying vertical position of the tip, hence of the surface contour, as a function of the tip position in the xy plane.

An STM not only can provide an image of a static surface, it can also be used to manipulate atoms and molecules on a surface, such as was done in forming the *quantum corral* shown in Fig. 39-12 in the next chapter. In a process known as lateral manipulation, the STM probe is initially brought down near a molecule, close enough that the molecule is attracted to the probe without actually touching it. The probe is then moved across the background surface (such as platinum), dragging the molecule with it until the molecule is in the desired location. Then the probe is backed up away from the molecule, weakening and then eliminating the attractive force on the molecule. Although the work requires very fine control, a design can eventually be formed. In Fig. 39-12, an STM probe has been used to move 48 iron atoms across a copper surface and into a circular corral 14 nm in diameter, in which electrons can be trapped.

Sample Problem

Barrier tunneling by matter wave

Suppose that the electron in Fig. 38-15, having a total energy E of 5.1 eV, approaches a barrier of height $U_b = 6.8$ eV and thickness $L = 750$ pm.

- (a) What is the approximate probability that the electron will be transmitted through the barrier, to appear (and be detectable) on the other side of the barrier?

KEY IDEA

The probability we seek is the transmission coefficient T as given by Eq. 38-21 ($T \approx e^{-2bL}$), where

$$b = \sqrt{\frac{8\pi^2 m(U_b - E)}{h^2}}$$

Calculations: The numerator of the fraction under the square-root sign is

$$(8\pi^2)(9.11 \times 10^{-31} \text{ kg})(6.8 \text{ eV} - 5.1 \text{ eV}) \\ \times (1.60 \times 10^{-19} \text{ J/eV}) = 1.956 \times 10^{-47} \text{ J} \cdot \text{kg}$$

$$\text{Thus, } b = \sqrt{\frac{1.956 \times 10^{-47} \text{ J} \cdot \text{kg}}{(6.63 \times 10^{-34} \text{ J} \cdot \text{s})^2}} = 6.67 \times 10^9 \text{ m}^{-1}$$

The (dimensionless) quantity $2bL$ is then

$$2bL = (2)(6.67 \times 10^9 \text{ m}^{-1})(750 \times 10^{-12} \text{ m}) = 10.0$$

and, from Eq. 38-21, the transmission coefficient is

$$T \approx e^{-2bL} = e^{-10.0} = 45 \times 10^{-6} \quad (\text{Answer})$$

Thus, of every million electrons that strike the barrier, about 45 will tunnel through it, each appearing on the other side with its original total energy of 5.1 eV. (The transmission through the barrier does not alter an electron's energy or any other property.)

- (b) What is the approximate probability that a proton with the same total energy of 5.1 eV will be transmitted through the barrier, to appear (and be detectable) on the other side of the barrier?

Reasoning: The transmission coefficient T (and thus the probability of transmission) depends on the mass of the particle. Indeed, because mass m is one of the factors in the exponent of e in the equation for T , the probability of transmission is very sensitive to the mass of the particle. This time, the mass is that of a proton (1.67×10^{-27} kg), which is significantly greater than that of the electron in (a). By substituting the proton's mass for the mass in (a) and then continuing as we did there, we find that $T \approx 10^{-186}$. Thus, although the probability that the proton will be transmitted is not exactly zero, it is barely more than zero. For even more massive particles with the same total energy of 5.1 eV, the probability of transmission is exponentially lower.



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REVIEW & SUMMARY

Light Quanta–Photons An electromagnetic wave (light) is quantized, and its quanta are called *photons*. For a light wave of frequency f and wavelength λ , the energy E and momentum magnitude p of a photon are

$$E = hf \quad (\text{photon energy}) \quad (38-2)$$

and $p = \frac{hf}{c} = \frac{h}{\lambda}$ (photon momentum). $(38-7)$

Photoelectric Effect When light of high enough frequency falls on a clean metal surface, electrons are emitted from the surface by photon–electron interactions within the metal. The governing relation is

$$hf = K_{\max} + \Phi, \quad (38-5)$$

in which hf is the photon energy, K_{\max} is the kinetic energy of the most energetic emitted electrons, and Φ is the **work function** of the target material—that is, the minimum energy an electron must have if it is to emerge from the surface of the target. If hf is less than Φ , electrons are not emitted.

Compton Shift When x rays are scattered by loosely bound electrons in a target, some of the scattered x rays have a longer wavelength than do the incident x rays. This **Compton shift** (in wavelength) is given by

$$\Delta\lambda = \frac{h}{mc}(1 - \cos\phi), \quad (38-11)$$

in which ϕ is the angle at which the x rays are scattered.

Light Waves and Photons When light interacts with matter, energy and momentum are transferred via photons. When light is in transit, however, we interpret the light wave as a **probability wave**, in which the probability (per unit time) that a photon can be detected is proportional to E_m^2 , where E_m is the amplitude of the oscillating electric field of the light wave at the detector.

Matter Waves A moving particle such as an electron or a proton can be described as a **matter wave**; its wavelength (called the **de Broglie wavelength**) is given by $\lambda = h/p$, where p is the magnitude of the particle's momentum.

The Wave Function A matter wave is described by its **wave function** $\Psi(x, y, z, t)$, which can be separated into a space-dependent part $\psi(x, y, z)$ and a time-dependent part $e^{-i\omega t}$. For a particle of mass m moving in the x direction with constant total energy E through a region in which its potential energy is $U(x)$, $\psi(x)$ can be found by solving the simplified **Schrödinger equation**:

$$\frac{d^2\psi}{dx^2} + \frac{8\pi^2m}{h^2}[E - U(x)]\psi = 0. \quad (38-15)$$

A matter wave, like a light wave, is a probability wave in the sense that if a particle detector is inserted into the wave, the probability that the detector will register a particle during any specified time interval is proportional to $|\psi|^2$, a quantity called the **probability density**.

For a free particle—that is, a particle for which $U(x) = 0$ —moving in the x direction, $|\psi|^2$ has a constant value for all positions along the x axis.

Heisenberg's Uncertainty Principle The probabilistic nature of quantum physics places an important limitation on detecting a particle's position and momentum. That is, it is not possible to measure the position \vec{r} and the momentum \vec{p} of a particle simultaneously with unlimited precision. The uncertainties in the components of these quantities are given by

$$\begin{aligned} \Delta x \cdot \Delta p_x &\geq \hbar \\ \Delta y \cdot \Delta p_y &\geq \hbar \\ \Delta z \cdot \Delta p_z &\geq \hbar. \end{aligned} \quad (38-20)$$

Barrier Tunneling According to classical physics, an incident particle will be reflected from a potential energy barrier whose height is greater than the particle's kinetic energy. According to quantum physics, however, the particle has a finite probability of tunneling through such a barrier. The probability that a given particle of mass m and energy E will tunnel through a barrier of height U_b and thickness L is given by the transmission coefficient T :

$$T \approx e^{-2bL}, \quad (38-21)$$

where $b = \sqrt{\frac{8\pi^2m(U_b - E)}{h^2}}$. $(38-22)$

QUESTIONS

1 Photon A has twice the energy of photon B. (a) Is the momentum of A less than, equal to, or greater than that of B? (b) Is the wavelength of A less than, equal to, or greater than that of B?

2 In the photoelectric effect (for a given target and a given frequency of the incident light), which of these quantities, if any, depend on the intensity of the incident light beam: (a) the maximum kinetic energy of the electrons, (b) the maximum photoelectric current, (c) the stopping potential, (d) the cutoff frequency?

3 According to the figure for Checkpoint 2, is the maximum kinetic energy of the ejected electrons greater for a target made of sodium or of potassium for a given frequency of incident light?

4 Photoelectric effect: Figure 38-18 gives the stopping voltage V

versus the wavelength λ of light for three different materials. Rank the materials according to their work function, greatest first.

5 A metal plate is illuminated with light of a certain frequency. Which of the following determine whether or not electrons are ejected: (a) the intensity of the light, (b) how long the plate is exposed to the light, (c) the thermal conductivity of the plate, (d) the area of the plate, (e) the material of which the plate is made?

6 Let K be the kinetic energy that a stationary free electron gains

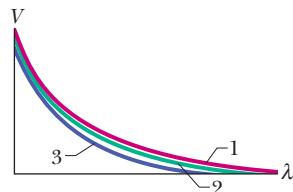


Fig. 38-18 Question 4.

1078 CHAPTER 38 PHOTONS AND MATTER WAVES

when a photon scatters from it. We can plot K versus the angle ϕ at which the photon scatters; see curve 1 in Fig. 38-19. If we switch the target to be a stationary free proton, does the end point of the graph shift (a) upward as suggested by curve 2, (b) downward as suggested by curve 3, or (c) remain the same?

- 7** In a Compton-shift experiment, light (in the x-ray range) is scattered in the forward direction, at $\phi = 0$ in Fig. 38-3. What fraction of the light's energy does the electron acquire?

- 8** *Compton scattering.* Figure 38-20 gives the Compton shift $\Delta\lambda$ versus scattering angle ϕ for three different stationary target particles. Rank the particles according to their mass, greatest first.

- 9** (a) If you double the kinetic energy of a nonrelativistic particle, how does its de Broglie wavelength change? (b) What if you double the speed of the particle?

- 10** Figure 38-21 shows an electron moving (a) opposite an electric field, (b) in the same direction as an electric field, (c) in the same direction as a magnetic field, (d) perpendicular to a magnetic field. For each situation, is the de Broglie wavelength of the electron increasing, decreasing, or remaining the same?

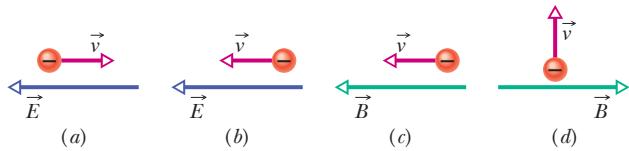


Fig. 38-21 Question 10.

- 11** At the left in Fig. 38-16, why are the minima nonzero?

- 12** An electron and a proton have the same kinetic energy. Which has the greater de Broglie wavelength?

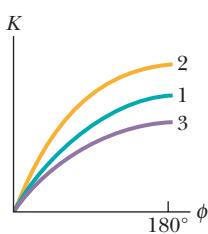


Fig. 38-19
Question 6.

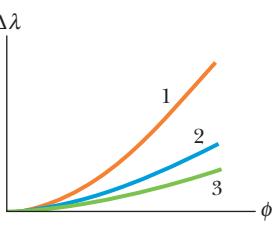


Fig. 38-20 Question 8.

- 13** The following nonrelativistic particles all have the same kinetic energy. Rank them in order of their de Broglie wavelengths, greatest first: electron, alpha particle, neutron.

- 14** Figure 38-22 shows an electron moving through several regions where uniform electric potentials V have been set up. Rank the three regions according to the de Broglie wavelength of the electron there, greatest first.

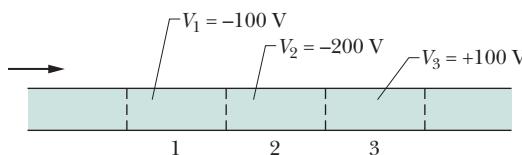


Fig. 38-22 Question 14.

- 15** The table gives relative values for three situations for the barrier tunneling experiment of Figs. 38-14 and 38-15. Rank the situations according to the probability of the electron tunneling through the barrier, greatest first.

	Electron Energy	Barrier Height	Barrier Thickness
(a)	E	$5E$	L
(b)	E	$17E$	$L/2$
(c)	E	$2E$	$2L$

- 16** For three experiments, Fig. 38-23 gives the transmission coefficient T for electron tunneling through a potential barrier, plotted versus barrier thickness L . The de Broglie wavelengths of the electrons are identical in the three experiments. The only difference in the physical setups is the barrier heights U_b . Rank the three experiments according to U_b , greatest first.

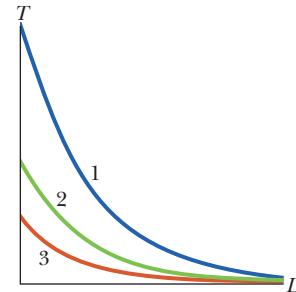


Fig. 38-23 Question 16.

P R O B L E M S



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>



sec. 38-2 The Photon, the Quantum of Light

- 1** Monochromatic light (that is, light of a single wavelength) is to be absorbed by a sheet of photographic film and thus recorded on the film. Photon absorption will occur if the photon energy equals or exceeds 0.6 eV, the smallest amount of energy needed to dissociate an AgBr molecule in the film. (a) What is the greatest wavelength of light that can be recorded by the film? (b) In what region of the electromagnetic spectrum is this wavelength located?

- 2** How fast must an electron move to have a kinetic energy equal to the photon energy of sodium light at wavelength 590 nm?

- 3** At what rate does the Sun emit photons? For simplicity, assume that the Sun's entire emission at the rate of 3.9×10^{26} W is at the single wavelength of 550 nm.

- 4** A helium-neon laser emits red light at wavelength $\lambda = 633$ nm in a beam of diameter 3.5 mm and at an energy-emission rate of 5.0 mW. A detector in the beam's path totally absorbs the beam. At what rate per unit area does the detector absorb photons?

- 5** The meter was once defined as 1 650 763.73 wavelengths of the orange light emitted by a source containing krypton-86 atoms. What is the photon energy of that light?

•6 The yellow-colored light from a highway sodium lamp is brightest at a wavelength of 589 nm. What is the photon energy for light at that wavelength?

•7 A light detector (your eye) has an area of $2.00 \times 10^{-6} \text{ m}^2$ and absorbs 80% of the incident light, which is at wavelength 500 nm. The detector faces an isotropic source, 3.00 m from the source. If the detector absorbs photons at the rate of exactly 4.000 s^{-1} , at what power does the emitter emit light?

•8 The beam emerging from a 1.5 W argon laser ($\lambda = 515 \text{ nm}$) has a diameter d of 3.0 mm. The beam is focused by a lens system with an effective focal length f_L of 2.5 mm. The focused beam strikes a totally absorbing screen, where it forms a circular diffraction pattern whose central disk has a radius R given by $1.22f_L\lambda/d$. It can be shown that 84% of the incident energy ends up within this central disk. At what rate are photons absorbed by the screen in the central disk of the diffraction pattern?

•9 A 100 W sodium lamp ($\lambda = 589 \text{ nm}$) radiates energy uniformly in all directions. (a) At what rate are photons emitted by the lamp? (b) At what distance from the lamp will a totally absorbing screen absorb photons at the rate of $1.00 \text{ photon/cm}^2 \cdot \text{s}$? (c) What is the photon flux (photons per unit area per unit time) on a small screen 2.00 m from the lamp?

•10 A satellite in Earth orbit maintains a panel of solar cells of area 2.60 m^2 perpendicular to the direction of the Sun's light rays. The intensity of the light at the panel is 1.39 kW/m^2 . (a) At what rate does solar energy arrive at the panel? (b) At what rate are solar photons absorbed by the panel? Assume that the solar radiation is monochromatic, with a wavelength of 550 nm, and that all the solar radiation striking the panel is absorbed. (c) How long would it take for a "mole of photons" to be absorbed by the panel?

•11 An ultraviolet lamp emits light of wavelength 400 nm at the rate of 400 W. An infrared lamp emits light of wavelength 700 nm, also at the rate of 400 W. (a) Which lamp emits photons at the greater rate and (b) what is that greater rate?

•12 Under ideal conditions, a visual sensation can occur in the human visual system if light of wavelength 550 nm is absorbed by the eye's retina at a rate as low as 100 photons per second. What is the corresponding rate at which energy is absorbed by the retina?

•13 A special kind of lightbulb emits monochromatic light of wavelength 630 nm. Electrical energy is supplied to it at the rate of 60 W, and the bulb is 93% efficient at converting that energy to light energy. How many photons are emitted by the bulb during its lifetime of 730 h?

•14 A light detector has an absorbing area of $2.00 \times 10^{-6} \text{ m}^2$ and absorbs 50% of the incident light, which is at wavelength 600 nm. The detector faces an isotropic source, 12.0 m from the source. The energy E emitted by the source versus time t is given in Fig. 38-24 ($E_s = 7.2 \text{ nJ}$, $t_s = 2.0 \text{ s}$). At what rate are photons absorbed by the detector?

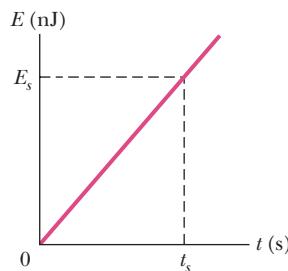


Fig. 38-24 Problem 14.

sec. 38-3 The Photoelectric Effect

•15 Light strikes a sodium surface, causing photoelectric emission. The stopping potential for the ejected electrons is 5.0 V,

and the work function of sodium is 2.2 eV. What is the wavelength of the incident light?

•16 Find the maximum kinetic energy of electrons ejected from a certain material if the material's work function is 2.3 eV and the frequency of the incident radiation is $3.0 \times 10^{15} \text{ Hz}$.

•17 The work function of tungsten is 4.50 eV. Calculate the speed of the fastest electrons ejected from a tungsten surface when light whose photon energy is 5.80 eV shines on the surface.

•18 You wish to pick an element for a photocell that will operate via the photoelectric effect with visible light. Which of the following are suitable (work functions are in parentheses): tantalum (4.2 eV), tungsten (4.5 eV), aluminum (4.2 eV), barium (2.5 eV), lithium (2.3 eV)?

•19 (a) If the work function for a certain metal is 1.8 eV, what is the stopping potential for electrons ejected from the metal when light of wavelength 400 nm shines on the metal? (b) What is the maximum speed of the ejected electrons?

•20 Suppose the *fractional efficiency* of a cesium surface (with work function 1.80 eV) is 1.0×10^{-16} ; that is, on average one electron is ejected for every 10^{16} photons that reach the surface. What would be the current of electrons ejected from such a surface if it were illuminated with 600 nm light from a 2.00 mW laser and all the ejected electrons took part in the charge flow?

•21 X rays with a wavelength of 71 pm are directed onto a gold foil and eject tightly bound electrons from the gold atoms. The ejected electrons then move in circular paths of radius r in a region of uniform magnetic field \vec{B} . For the fastest of the ejected electrons, the product Br is equal to $1.88 \times 10^{-4} \text{ T} \cdot \text{m}$. Find (a) the maximum kinetic energy of those electrons and (b) the work done in removing them from the gold atoms.

•22 The wavelength associated with the cutoff frequency for silver is 325 nm. Find the maximum kinetic energy of electrons ejected from a silver surface by ultraviolet light of wavelength 254 nm.

•23 Light of wavelength 200 nm shines on an aluminum surface; 4.20 eV is required to eject an electron. What is the kinetic energy of (a) the fastest and (b) the slowest ejected electrons? (c) What is the stopping potential for this situation? (d) What is the cutoff wavelength for aluminum?

•24 In a photoelectric experiment using a sodium surface, you find a stopping potential of 1.85 V for a wavelength of 300 nm and a stopping potential of 0.820 V for a wavelength of 400 nm. From these data find (a) a value for the Planck constant, (b) the work function Φ for sodium, and (c) the cutoff wavelength λ_0 for sodium.

•25 The stopping potential for electrons emitted from a surface illuminated by light of wavelength 491 nm is 0.710 V. When the incident wavelength is changed to a new value, the stopping potential is 1.43 V. (a) What is this new wavelength? (b) What is the work function for the surface?

•26 An orbiting satellite can become charged by the photoelectric effect when sunlight ejects electrons from its outer surface. Satellites must be designed to minimize such charging because it can ruin the sensitive microelectronics. Suppose a satellite is coated with platinum, a metal with a very large work function ($\Phi = 5.32 \text{ eV}$). Find the longest wavelength of incident sunlight that can eject an electron from the platinum.

1080 CHAPTER 38 PHOTONS AND MATTER WAVES

sec. 38-4 Photons Have Momentum

•27 **SSM** Light of wavelength 2.40 pm is directed onto a target containing free electrons. (a) Find the wavelength of light scattered at 30.0° from the incident direction. (b) Do the same for a scattering angle of 120° .

•28 (a) In MeV/c , what is the magnitude of the momentum associated with a photon having an energy equal to the electron rest energy? What are the (b) wavelength and (c) frequency of the corresponding radiation?

•29 What (a) frequency, (b) photon energy, and (c) photon momentum magnitude (in keV/c) are associated with x rays having wavelength 35.0 pm?

•30 What is the maximum wavelength shift for a Compton collision between a photon and a free *proton*?

•31 What percentage increase in wavelength leads to a 75% loss of photon energy in a photon–free electron collision?

•32 X rays of wavelength 0.0100 nm are directed in the positive direction of an *x* axis onto a target containing loosely bound electrons. For Compton scattering from one of those electrons, at an angle of 180° , what are (a) the Compton shift, (b) the corresponding change in photon energy, (c) the kinetic energy of the recoiling electron, and (d) the angle between the positive direction of the *x* axis and the electron’s direction of motion?

•33 Calculate the percentage change in photon energy during a collision like that in Fig. 38-5 for $\phi = 90^\circ$ and for radiation in (a) the microwave range, with $\lambda = 3.0 \text{ cm}$; (b) the visible range, with $\lambda = 500 \text{ nm}$; (c) the x-ray range, with $\lambda = 25 \text{ pm}$; and (d) the gamma-ray range, with a gamma photon energy of 1.0 MeV. (e) What are your conclusions about the feasibility of detecting the Compton shift in these various regions of the electromagnetic spectrum, judging solely by the criterion of energy loss in a single photon–electron encounter?

•34 A photon undergoes Compton scattering off a stationary free electron. The photon scatters at 90.0° from its initial direction; its initial wavelength is $3.00 \times 10^{-12} \text{ m}$. What is the electron’s kinetic energy?

•35 Calculate the Compton wavelength for (a) an electron and (b) a proton. What is the photon energy for an electromagnetic wave with a wavelength equal to the Compton wavelength of (c) the electron and (d) the proton?

•36 Gamma rays of photon energy 0.511 MeV are directed onto an aluminum target and are scattered in various directions by loosely bound electrons there. (a) What is the wavelength of the incident gamma rays? (b) What is the wavelength of gamma rays scattered at 90.0° to the incident beam? (c) What is the photon energy of the rays scattered in this direction?

•37 Consider a collision between an x-ray photon of initial energy 50.0 keV and an electron at rest, in which the photon is scattered backward and the electron is knocked forward. (a) What is the energy of the back-scattered photon? (b) What is the kinetic energy of the electron?

•38 Show that when a photon of energy E is scattered from a free electron at rest, the maximum kinetic energy of the recoiling electron is given by

$$K_{\max} = \frac{E^2}{E + mc^2/2}.$$

•39 Through what angle must a 200 keV photon be scattered by a free electron so that the photon loses 10% of its energy?

•40 What is the maximum kinetic energy of electrons knocked out of a thin copper foil by Compton scattering of an incident beam of 17.5 keV x rays? Assume the work function is negligible.

•41 What are (a) the Compton shift $\Delta\lambda$, (b) the fractional Compton shift $\Delta\lambda/\lambda$, and (c) the change ΔE in photon energy for light of wavelength $\lambda = 590 \text{ nm}$ scattering from a free, initially stationary electron if the scattering is at 90° to the direction of the incident beam? What are (d) $\Delta\lambda$, (e) $\Delta\lambda/\lambda$, and (f) ΔE for 90° scattering for photon energy 50.0 keV (x-ray range)?

sec. 38-6 Electrons and Matter Waves

•42 Calculate the de Broglie wavelength of (a) a 1.00 keV electron, (b) a 1.00 keV photon, and (c) a 1.00 keV neutron.

•43 **SSM** In an old-fashioned television set, electrons are accelerated through a potential difference of 25.0 kV. What is the de Broglie wavelength of such electrons? (Relativity is not needed.)

•44 The smallest dimension (*resolving power*) that can be resolved by an electron microscope is equal to the de Broglie wavelength of its electrons. What accelerating voltage would be required for the electrons to have the same resolving power as could be obtained using 100 keV gamma rays?

•45 **SSM** Singly charged sodium ions are accelerated through a potential difference of 300 V. (a) What is the momentum acquired by such an ion? (b) What is its de Broglie wavelength?

•46 Electrons accelerated to an energy of 50 GeV have a de Broglie wavelength λ small enough for them to probe the structure within a target nucleus by scattering from the structure. Assume that the energy is so large that the extreme relativistic relation $p = E/c$ between momentum magnitude p and energy E applies. (In this extreme situation, the kinetic energy of an electron is much greater than its rest energy.) (a) What is λ ? (b) If the target nucleus has radius $R = 5.0 \text{ fm}$, what is the ratio R/λ ?

•47 **SSM** The wavelength of the yellow spectral emission line of sodium is 590 nm. At what kinetic energy would an electron have that wavelength as its de Broglie wavelength?

•48 A stream of protons, each with a speed of $0.9900c$, are directed into a two-slit experiment where the slit separation is $4.00 \times 10^{-9} \text{ m}$. A two-slit interference pattern is built up on the viewing screen. What is the angle between the center of the pattern and the second minimum (to either side of the center)?

•49 What is the wavelength of (a) a photon with energy 1.00 eV, (b) an electron with energy 1.00 eV, (c) a photon of energy 1.00 GeV, and (d) an electron with energy 1.00 GeV?

•50 An electron and a photon each have a wavelength of 0.20 nm. What is the momentum (in $\text{kg} \cdot \text{m/s}$) of the (a) electron and (b) photon? What is the energy (in eV) of the (c) electron and (d) photon?

•51 The highest achievable resolving power of a microscope is limited only by the wavelength used; that is, the smallest item that can be distinguished has dimensions about equal to the wavelength. Suppose one wishes to “see” inside an atom. Assuming the atom to have a diameter of 100 pm, this means that one must be able to resolve a width of, say, 10 pm. (a) If an electron microscope is used, what minimum electron energy is required? (b) If a light microscope is used, what minimum photon energy is required? (c) Which microscope seems more practical? Why?

••52 The existence of the atomic nucleus was discovered in 1911 by Ernest Rutherford, who properly interpreted some experiments in which a beam of alpha particles was scattered from a metal foil of atoms such as gold. (a) If the alpha particles had a kinetic energy of 7.5 MeV, what was their de Broglie wavelength? (b) Explain whether the wave nature of the incident alpha particles should have been taken into account in interpreting these experiments. The mass of an alpha particle is 4.00 u (atomic mass units), and its distance of closest approach to the nuclear center in these experiments was about 30 fm. (The wave nature of matter was not postulated until more than a decade after these crucial experiments were first performed.)

••53 A nonrelativistic particle is moving three times as fast as an electron. The ratio of the de Broglie wavelength of the particle to that of the electron is 1.813×10^{-4} . By calculating its mass, identify the particle.

••54 What are (a) the energy of a photon corresponding to wavelength 1.00 nm, (b) the kinetic energy of an electron with de Broglie wavelength 1.00 nm, (c) the energy of a photon corresponding to wavelength 1.00 fm, and (d) the kinetic energy of an electron with de Broglie wavelength 1.00 fm?

••55 If the de Broglie wavelength of a proton is 100 fm, (a) what is the speed of the proton and (b) through what electric potential would the proton have to be accelerated to acquire this speed?

sec. 38-7 Schrödinger's Equation

••56 Suppose we put $A = 0$ in Eq. 38-17 and relabeled B as ψ_0 . (a) What would the resulting wave function then describe? (b) How, if at all, would Fig. 38-12 be altered?

••57 The function $\psi(x)$ displayed in Eq. 38-19 describes a free particle, for which we assumed that $U(x) = 0$ in Schrödinger's equation (Eq. 38-15). Assume now that $U(x) = U_0$ = a constant in that equation. Show that Eq. 38-19 is still a solution of Schrödinger's equation, with

$$k = \frac{2\pi}{h} \sqrt{2m(E - U_0)}$$

now giving the angular wave number k of the particle.

••58 In Eq. 38-18 keep both terms, putting $A = B = \psi_0$. The equation then describes the superposition of two matter waves of equal amplitude, traveling in opposite directions. (Recall that this is the condition for a standing wave.) (a) Show that $|\Psi(x, t)|^2$ is then given by

$$|\Psi(x, t)|^2 = 2\psi_0^2[1 + \cos 2kx].$$

(b) Plot this function, and demonstrate that it describes the square of the amplitude of a standing matter wave. (c) Show that the nodes of this standing wave are located at

$$x = (2n + 1)\left(\frac{1}{4}\lambda\right), \quad \text{where } n = 0, 1, 2, 3, \dots$$

and λ is the de Broglie wavelength of the particle. (d) Write a similar expression for the most probable locations of the particle.

••59 Show that Eq. 38-17 is indeed a solution of Eq. 38-16 by substituting $\psi(x)$ and its second derivative into Eq. 38-16 and noting that an identity results.

••60 (a) Write the wave function $\psi(x)$ displayed in Eq. 38-19 in the form $\psi(x) = a + ib$, where a and b are real quantities. (Assume that ψ_0 is real.) (b) Write the time-dependent wave function $\Psi(x, t)$ that corresponds to $\psi(x)$ written in this form.

••61 Show that the angular wave number k for a nonrelativistic free particle of mass m can be written as

$$k = \frac{2\pi \sqrt{2mK}}{h},$$

in which K is the particle's kinetic energy.

••62 (a) Let $n = a + ib$ be a complex number, where a and b are real (positive or negative) numbers. Show that the product nn^* is always a positive real number. (b) Let $m = c + id$ be another complex number. Show that $|nm| = |n||m|$.

sec. 38-8 Heisenberg's Uncertainty Principle

••63 The uncertainty in the position of an electron along an x axis is given as 50 pm, which is about equal to the radius of a hydrogen atom. What is the least uncertainty in any simultaneous measurement of the momentum component p_x of this electron?

••64 You will find in Chapter 39 that electrons cannot move in definite orbits within atoms, like the planets in our solar system. To see why, let us try to "observe" such an orbiting electron by using a light microscope to measure the electron's presumed orbital position with a precision of, say, 10 pm (a typical atom has a radius of about 100 pm). The wavelength of the light used in the microscope must then be about 10 pm. (a) What would be the photon energy of this light? (b) How much energy would such a photon impart to an electron in a head-on collision? (c) What do these results tell you about the possibility of "viewing" an atomic electron at two or more points along its presumed orbital path? (Hint: The outer electrons of atoms are bound to the atom by energies of only a few electron-volts.)

••65 Figure 38-12 shows a case in which the momentum component p_x of a particle is fixed so that $\Delta p_x = 0$; then, from Heisenberg's uncertainty principle (Eq. 38-20), the position x of the particle is completely unknown. From the same principle it follows that the opposite is also true; that is, if the position of a particle is exactly known ($\Delta x = 0$), the uncertainty in its momentum is infinite.

Consider an intermediate case, in which the position of a particle is measured, not to infinite precision, but to within a distance of $\lambda/2\pi$, where λ is the particle's de Broglie wavelength. Show that the uncertainty in the (simultaneously measured) momentum component is then equal to the component itself; that is, $\Delta p_x = p$. Under these circumstances, would a measured momentum of zero surprise you? What about a measured momentum of $0.5p$? Of $2p$? Of $12p$?

sec. 38-9 Barrier Tunneling

••66 Consider a potential energy barrier like that of Fig. 38-15 but whose height U_b is 6.0 eV and whose thickness L is 0.70 nm. What is the energy of an incident electron whose transmission coefficient is 0.0010?

••67 A 3.0 MeV proton is incident on a potential energy barrier of thickness 10 fm and height 10 MeV. What are (a) the transmission coefficient T , (b) the kinetic energy K_t the proton will have on the other side of the barrier if it tunnels through the barrier, and (c) the kinetic energy K_r it will have if it reflects from the barrier? A 3.0 MeV deuteron (the same charge but twice the mass as a proton) is incident on the same barrier. What are (d) T , (e) K_t , and (f) K_r ?

••68 (a) Suppose a beam of 5.0 eV protons strikes a potential energy barrier of height 6.0 eV and thickness 0.70 nm, at a rate

1082 CHAPTER 38 PHOTONS AND MATTER WAVES

equivalent to a current of 1000 A. How long would you have to wait—on average—for one proton to be transmitted? (b) How long would you have to wait if the beam consisted of electrons rather than protons?

•69 SSM WWW An electron with total energy $E = 5.1$ eV approaches a barrier of height $U_b = 6.8$ eV and thickness $L = 750$ pm. What percentage change in the transmission coefficient T occurs for a 1.0% change in (a) the barrier height, (b) the barrier thickness, and (c) the kinetic energy of the incident electron?

Additional Problems

70 Figure 38-12 shows that because of Heisenberg's uncertainty principle, it is not possible to assign an x coordinate to the position of a free electron moving along an x axis. (a) Can you assign a y or a z coordinate? (*Hint:* The momentum of the electron has no y or z component.) (b) Describe the extent of the matter wave in three dimensions.

71 A spectral emission line is electromagnetic radiation that is emitted in a wavelength range narrow enough to be taken as a single wavelength. One such emission line that is important in astronomy has a wavelength of 21 cm. What is the photon energy in the electromagnetic wave at that wavelength?

72 Using the classical equations for momentum and kinetic energy, show that an electron's de Broglie wavelength in nanometers can be written as $\lambda = 1.226/\sqrt{K}$, in which K is the electron's kinetic energy in electron-volts.

73 Derive Eq. 38-11, the equation for the Compton shift, from Eqs. 38-8, 38-9, and 38-10 by eliminating v and θ .

74 Neutrons in thermal equilibrium with matter have an average kinetic energy of $(3/2)kT$, where k is the Boltzmann constant and T , which may be taken to be 300 K, is the temperature of the environment of the neutrons. (a) What is the average kinetic energy of such a neutron? (b) What is the corresponding de Broglie wavelength?

75 Consider a balloon filled with helium gas at room temperature and atmospheric pressure. Calculate (a) the average de Broglie wavelength of the helium atoms and (b) the average distance between atoms under these conditions. The average kinetic energy of an atom is equal to $(3/2)kT$, where k is the Boltzmann constant. (c) Can the atoms be treated as particles under these conditions? Explain.

76 In about 1916, R. A. Millikan found the following stopping-potential data for lithium in his photoelectric experiments:

Wavelength (nm)	433.9	404.7	365.0	312.5	253.5
Stopping potential (V)	0.55	0.73	1.09	1.67	2.57

Use these data to make a plot like Fig. 38-2 (which is for sodium) and then use the plot to find (a) the Planck constant and (b) the work function for lithium.

77 Show that $|\psi|^2 = |\Psi|^2$, with ψ and Ψ related as in Eq. 38-14. That is, show that the probability density does not depend on the time variable.

78 Show that $\Delta E/E$, the fractional loss of energy of a photon during a collision with a particle of mass m , is given by

$$\frac{\Delta E}{E} = \frac{hf'}{mc^2} (1 - \cos \phi),$$

where E is the energy of the incident photon, f' is the frequency of the scattered photon, and ϕ is defined as in Fig. 38-5.

79 A bullet of mass 40 g travels at 1000 m/s. Although the bullet is clearly too large to be treated as a matter wave, determine what Eq. 38-13 predicts for the de Broglie wavelength of the bullet at that speed.

80 (a) The smallest amount of energy needed to eject an electron from metallic sodium is 2.28 eV. Does sodium show a photoelectric effect for red light, with $\lambda = 680$ nm? (That is, does the light cause electron emission?) (b) What is the cutoff wavelength for photoelectric emission from sodium? (c) To what color does that wavelength correspond?

81 SSM Imagine playing baseball in a universe (not ours!) where the Planck constant is 0.60 J·s and thus quantum physics affects macroscopic objects. What would be the uncertainty in the position of a 0.50 kg baseball that is moving at 20 m/s along an axis if the uncertainty in the speed is 1.0 m/s?

82 An electron of mass m and speed v “collides” with a gamma-ray photon of initial energy hf_0 , as measured in the laboratory frame. The photon is scattered in the electron's direction of travel. Verify that the energy of the scattered photon, as measured in the laboratory frame, is

$$E = hf_0 \left(1 + \frac{2hf_0}{mc^2} \sqrt{\frac{1+v/c}{1-v/c}} \right)^{-1}.$$

83 Show, by analyzing a collision between a photon and a free electron (using relativistic mechanics), that it is impossible for a photon to transfer all its energy to a free electron (and thus for the photon to vanish).

84 A 1500 kg car moving at 20 m/s approaches a hill that is 24 m high and 30 m long. Although the car and hill are clearly too large to be treated as matter waves, determine what Eq. 38-21 predicts for the transmission coefficient of the car, as if it could tunnel through the hill as a matter wave. Treat the hill as a potential energy barrier where the potential energy is gravitational.

39

MORE ABOUT MATTER WAVES

39-1 WHAT IS PHYSICS?

One of the long-standing goals of physics has been to understand the nature of atoms. Early in the 20th century nobody knew how the electrons in an atom are arranged, what their motions are, how atoms emit or absorb light, or even why atoms are stable. Without this knowledge it was not possible to understand how atoms combine to form molecules or stack up to form solids. As a consequence, the foundations of chemistry—including biochemistry, which underlies the nature of life itself—were more or less a mystery.

In 1926, all these questions and many others were answered with the development of quantum physics. Its basic premise is that moving electrons, protons, and particles of any kind are best viewed as matter waves, whose motions are governed by Schrödinger's equation. Although quantum theory also applies to massive particles, there is no point in treating baseballs, automobiles, planets, and such objects with quantum theory. For such massive, slow-moving objects, Newtonian physics and quantum physics yield the same answers.

Before we can apply quantum physics to the problem of atomic structure, we need to develop some insights by applying quantum ideas in a few simpler situations. Some of these situations may seem simplistic and unreal, but they allow us to discuss the basic principles of the quantum physics of atoms without having to deal with the often overwhelming complexity of atoms. Besides, with the advances in nanotechnology, situations that were previously found only in textbooks are now being produced in laboratories and put to use in modern electronics and materials science applications. We are on the threshold of being able to use nanometer-scale constructions called *quantum corrals* and *quantum dots* to create “designer atoms” whose properties can be manipulated in the laboratory. For both natural atoms and these artificial ones, the starting point in our discussion is the wave nature of an electron.

39-2 String Waves and Matter Waves

In Chapter 16 we saw that waves of two kinds can be set up on a stretched string. If the string is so long that we can take it to be infinitely long, we can set up a *traveling wave* of essentially any frequency. However, if the stretched string has only a finite length, perhaps because it is rigidly clamped at both ends, we can set up only *standing waves* on it; further, these standing waves can have only discrete frequencies. In other words, confining the wave to a finite region of space leads to *quantization* of the motion—to the existence of discrete *states* for the wave, each state with a sharply defined frequency.

This observation applies to waves of all kinds, including matter waves. For matter waves, however, it is more convenient to deal with the energy E of the associated particle than with the frequency f of the wave. In all that follows we shall focus on the matter wave associated with an electron, but the results apply to any confined matter wave.

Consider the matter wave associated with an electron moving in the positive x direction and subject to no net force—a so-called *free particle*. The energy of such an electron can have any reasonable value, just as a wave traveling along a stretched string of infinite length can have any reasonable frequency.

Consider next the matter wave associated with an atomic electron, perhaps the *valence* (least tightly bound) electron. The electron—held within the atom by the attractive Coulomb force between it and the positively charged nucleus—is not a free particle. It can exist only in a set of discrete states, each having a discrete energy E . This sounds much like the discrete states and quantized frequencies that are available to a stretched string of finite length. For matter waves, then, as for all other kinds of waves, we may state a **confinement principle**:



Confinement of a wave leads to quantization—that is, to the existence of discrete states with discrete energies. The wave can have only those energies.

39-3 Energies of a Trapped Electron

One-Dimensional Traps

Here we examine the matter wave associated with a nonrelativistic electron confined to a limited region of space. We do so by analogy with standing waves on a string of finite length, stretched along an x axis and confined between rigid supports. Because the supports are rigid, the two ends of the string are nodes, or points at which the string is always at rest. There may be other nodes along the string, but these two must always be present, as Fig. 16-20 shows.

The states, or discrete standing wave patterns in which the string can oscillate, are those for which the length L of the string is equal to an integer number of half-wavelengths. That is, the string can occupy only states for which

$$L = \frac{n\lambda}{2}, \quad \text{for } n = 1, 2, 3, \dots \quad (39-1)$$

Each value of n identifies a state of the oscillating string; using the language of quantum physics, we can call the integer n a **quantum number**.

For each state of the string permitted by Eq. 39-1, the transverse displacement of the string at any position x along the string is given by

$$y_n(x) = A \sin\left(\frac{n\pi}{L}x\right), \quad \text{for } n = 1, 2, 3, \dots, \quad (39-2)$$

in which the quantum number n identifies the oscillation pattern and A depends on the time at which you inspect the string. (Equation 39-2 is a short version of Eq. 16-60.) We see that for all values of n and for all times, there is a point of zero displacement (a node) at $x = 0$ and at $x = L$, as there must be. Figure 16-19 shows time exposures of such a stretched string for $n = 2, 3$, and 4.

Now let us turn our attention to matter waves. Our first problem is to physically confine an electron that is moving along the x axis so that it remains within a finite segment of that axis. Figure 39-1 shows a conceivable one-dimensional *electron trap*. It consists of two semi-infinitely long cylinders, each of which has an electric potential approaching $-\infty$; between them is a hollow cylinder of length L , which has an electric potential of zero. We put a single electron into this central cylinder to trap it.

The trap of Fig. 39-1 is easy to analyze but is not very practical. Single electrons *can*, however, be trapped in the laboratory with traps that are more complex in design but similar in concept. At the University of Washington, for example, a single electron has been held in a trap for months on end, permitting scientists to make extremely precise measurements of its properties.

An electron can be trapped in the $V = 0$ region.

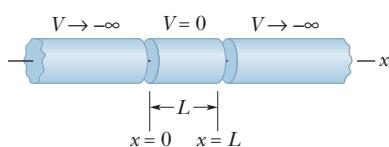


Fig. 39-1 The elements of an idealized “trap” designed to confine an electron to the central cylinder. We take the semi-infinitely long end cylinders to be at an infinitely great negative potential and the central cylinder to be at zero potential.

39-3 ENERGIES OF A TRAPPED ELECTRON

1085

Finding the Quantized Energies

Figure 39-2 shows the potential energy of the electron as a function of its position along the x axis of the idealized trap of Fig. 39-1. When the electron is in the central cylinder, its potential energy U ($= -eV$) is zero because there the potential V is zero. If the electron could get outside this region, its potential energy would be positive and of infinite magnitude because there $V \rightarrow -\infty$. We call the potential energy pattern of Fig. 39-2 an **infinitely deep potential energy well** or, for short, an *infinite potential well*. It is a “well” because an electron placed in the central cylinder of Fig. 39-1 cannot escape from it. As the electron approaches either end of the cylinder, a force of essentially infinite magnitude reverses the electron’s motion, thus trapping it. Because the electron can move along only a single axis, this trap can be called a *one-dimensional infinite potential well*.

Just like the standing wave in a length of stretched string, the matter wave describing the confined electron must have nodes at $x = 0$ and $x = L$. Moreover, Eq. 39-1 applies to such a matter wave if we interpret λ in that equation as the de Broglie wavelength associated with the moving electron.

The de Broglie wavelength λ is defined in Eq. 38-13 as $\lambda = h/p$, where p is the magnitude of the electron’s momentum. Because the electron is nonrelativistic, this momentum magnitude p is related to the kinetic energy K by $p = \sqrt{2mK}$, where m is the mass of the electron. For an electron moving within the central cylinder of Fig. 39-1, where $U = 0$, the total (mechanical) energy E is equal to the kinetic energy. Hence, we can write the de Broglie wavelength of this electron as

$$\lambda = \frac{h}{p} = \frac{h}{\sqrt{2mE}}. \quad (39-3)$$

If we substitute Eq. 39-3 into Eq. 39-1 and solve for the energy E , we find that E depends on n according to

$$E_n = \left(\frac{h^2}{8mL^2} \right) n^2, \quad \text{for } n = 1, 2, 3, \dots \quad (39-4)$$

The positive integer n here is the quantum number of the electron’s quantum state in the trap.

Equation 39-4 tells us something important: Because the electron is confined to the trap, it can have only the energies given by the equation. It *cannot* have an energy that is, say, halfway between the values for $n = 1$ and $n = 2$. Why this restriction? Because an electron is a matter wave. Were it, instead, a particle as assumed in classical physics, it could have *any* value of energy while it is confined to the trap.

Figure 39-3 is a graph showing the lowest five allowed energy values for an electron in an infinite well with $L = 100$ pm (about the size of a typical atom). The values are called *energy levels*, and they are drawn in Fig. 39-3 as levels, or steps, on a ladder, in an *energy-level diagram*. Energy is plotted vertically; nothing is plotted horizontally.

The quantum state with the lowest possible energy level E_1 allowed by Eq. 39-4, with quantum number $n = 1$, is called the *ground state* of the electron. The electron tends to be in this lowest energy state. All the quantum states with greater energies (corresponding to quantum numbers $n = 2$ or greater) are called *excited states* of the electron. The state with energy level E_2 , for quantum number $n = 2$, is called the *first excited state* because it is the first of the excited states as we move up the energy-level diagram. The other states have similar names.

Energy Changes

A trapped electron tends to have the lowest allowed energy and thus to be in its ground state. It can be changed to an excited state (in which it has greater energy)

An electron can be trapped in the $U = 0$ region.

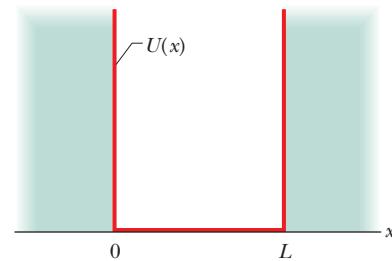


Fig. 39-2 The electric potential energy $U(x)$ of an electron confined to the central cylinder of the idealized trap of Fig. 39-1. We see that $U = 0$ for $0 < x < L$, and $U \rightarrow \infty$ for $x < 0$ and $x > L$.

These are the lowest five energy levels allowed the electron.
(No intermediate levels are allowed.)

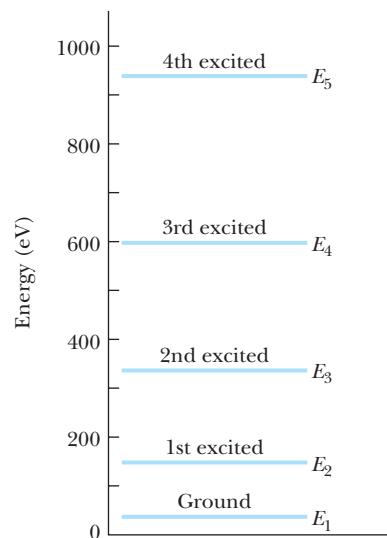
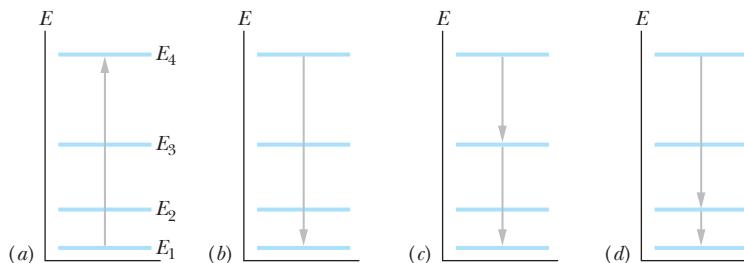


Fig. 39-3 Several of the allowed energies given by Eq. 39-4 for an electron confined to the infinite well of Fig. 39-2. Here width $L = 100$ pm. Such a plot is called an *energy-level diagram*.

Fig. 39-4 (a) Excitation of a trapped electron from the energy level of its ground state to the level of its third excited state. (b)–(d) Three of four possible ways the electron can de-excite to return to the energy level of its ground state. (Which way is not shown?)

The electron is excited to a higher energy level.
It can de-excite to a lower level in several ways (set by chance).



only if an external source provides the additional energy that is required for the change. Let E_{low} be the initial energy of the electron and E_{high} be the greater energy in a state that is higher on its energy-level diagram. Then the amount of energy that is required for the electron's change of state is

$$\Delta E = E_{\text{high}} - E_{\text{low}}. \quad (39-5)$$

An electron that receives such energy is said to make a *quantum jump* (or *transition*), or to be *excited* from the lower-energy state to the higher-energy state. Figure 39-4a represents a quantum jump from the ground state (with energy level E_1) to the third excited state (with energy level E_4). As shown, the jump *must* be from one energy level to another, but it can bypass one or more intermediate energy levels.

One way an electron can gain energy to make a quantum jump up to a greater energy level is to absorb a photon. However, this absorption and quantum jump can occur only if the following condition is met:

If a confined electron is to absorb a photon, the energy hf of the photon must equal the energy difference ΔE between the initial energy level of the electron and a higher level.

Thus, excitation by the absorption of light requires that

$$hf = \Delta E = E_{\text{high}} - E_{\text{low}}. \quad (39-6)$$

When an electron reaches an excited state, it does not stay there but quickly *de-excites* by decreasing its energy. Figures 39-4b to d represent some of the possible quantum jumps down from the energy level of the third excited state. The electron can reach its ground-state level either with one direct quantum jump (Fig. 39-4b) or with shorter jumps via intermediate levels (Figs. 39-4c and d).

An electron can decrease its energy by emitting a photon but only this way:

If a confined electron emits a photon, the energy hf of that photon must equal the energy difference ΔE between the initial energy level of the electron and a lower level.

Thus, Eq. 39-6 applies to both the absorption and the emission of light by a confined electron. That is, the absorbed or emitted light can have only certain values of hf and thus only certain values of frequency f and wavelength λ .

Aside: Although Eq. 39-6 and what we have discussed about photon absorption and emission can be applied to physical (real) electron traps, they actually cannot be applied to one-dimensional (unreal) electron traps. The reason involves the need to conserve angular momentum in a photon absorption or emission process. In this book, we shall neglect that need and use Eq. 39-6 even for one-dimensional traps.



CHECKPOINT 1

Rank the following pairs of quantum states for an electron confined to an infinite well according to the energy differences between the states, greatest first:
(a) $n = 3$ and $n = 1$, (b) $n = 5$ and $n = 4$,
(c) $n = 4$ and $n = 3$.

39-3 ENERGIES OF A TRAPPED ELECTRON

1087

Sample Problem

Energy levels in a 1D infinite potential well

An electron is confined to a one-dimensional, infinitely deep potential energy well of width $L = 100 \text{ pm}$.

- (a) What is the smallest amount of energy the electron can have?

KEY IDEA

Confinement of the electron (a matter wave) to the well leads to quantization of its energy. Because the well is infinitely deep, the allowed energies are given by Eq. 39-4 ($E_n = (h^2/8mL^2)n^2$), with the quantum number n a positive integer.

Lowest energy level: Here, the collection of constants in front of n^2 in Eq. 39-4 is evaluated as

$$\frac{h^2}{8mL^2} = \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})^2}{(8)(9.11 \times 10^{-31} \text{ kg})(100 \times 10^{-12} \text{ m})^2} = 6.031 \times 10^{-18} \text{ J}. \quad (39-7)$$

The smallest amount of energy the electron can have corresponds to the lowest quantum number, which is $n = 1$ for the ground state of the electron. Thus, Eqs. 39-4 and 39-7 give us

$$E_1 = \left(\frac{h^2}{8mL^2}\right)n^2 = (6.031 \times 10^{-18} \text{ J})(1^2) \approx 6.031 \times 10^{-18} \text{ J} = 37.7 \text{ eV}. \quad (\text{Answer})$$

- (b) How much energy must be transferred to the electron if it is to make a quantum jump from its ground state to its second excited state?

KEY IDEA

First a caution: Note that, from Fig. 39-3, the *second* excited state corresponds to the *third* energy level, with quantum number $n = 3$. Then if the electron is to jump from the $n = 1$ level to the $n = 3$ level, the required change in its energy is, from Eq. 39-5,

$$\Delta E_{31} = E_3 - E_1. \quad (39-8)$$

Upward jump: The energies E_3 and E_1 depend on the quantum number n , according to Eq. 39-4. Therefore, substituting that equation into Eq. 39-8 for energies E_3 and E_1 and

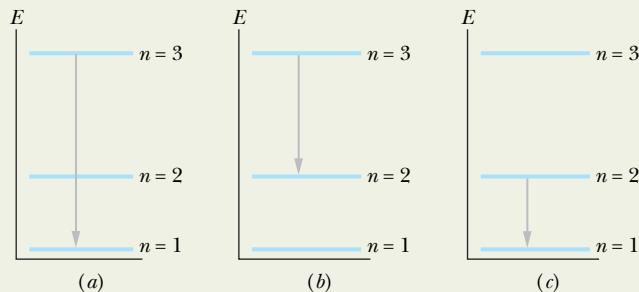


Fig. 39-5 De-excitation from the second excited state to the ground state either directly (a) or via the first excited state (b, c).

using Eq. 39-7 lead to

$$\begin{aligned} \Delta E_{31} &= \left(\frac{h^2}{8mL^2}\right)(3)^2 - \left(\frac{h^2}{8mL^2}\right)(1)^2 \\ &= \frac{h^2}{8mL^2}(3^2 - 1^2) \\ &= (6.031 \times 10^{-18} \text{ J})(8) \\ &= 4.83 \times 10^{-17} \text{ J} = 301 \text{ eV}. \quad (\text{Answer}) \end{aligned}$$

- (c) If the electron gains the energy for the jump from energy level E_1 to energy level E_3 by absorbing light, what light wavelength is required?

KEY IDEAS

(1) If light is to transfer energy to the electron, the transfer must be by photon absorption. (2) The photon's energy must equal the energy difference ΔE between the initial energy level of the electron and a higher level, according to Eq. 39-6 ($hf = \Delta E$). Otherwise, a photon *cannot* be absorbed.

Wavelength: Substituting c/λ for f , we can rewrite Eq. 39-6 as

$$\lambda = \frac{hc}{\Delta E}. \quad (39-9)$$

For the energy difference ΔE_{31} we found in (b), this equation gives us

$$\begin{aligned} \lambda &= \frac{hc}{\Delta E_{31}} \\ &= \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(2.998 \times 10^8 \text{ m/s})}{4.83 \times 10^{-17} \text{ J}} \\ &= 4.12 \times 10^{-9} \text{ m}. \quad (\text{Answer}) \end{aligned}$$

- (d) Once the electron has been excited to the second excited state, what wavelengths of light can it emit by de-excitation?

KEY IDEAS

1. The electron tends to de-excite, rather than remain in an excited state, until it reaches the ground state ($n = 1$).
2. If the electron is to de-excite, it must lose just enough energy to jump to a lower energy level.
3. If it is to lose energy by emitting light, then the loss of energy must be by emission of a photon.

Downward jumps: Starting in the second excited state (at the $n = 3$ level), the electron can reach the ground state ($n = 1$) by *either* making a quantum jump directly to the ground-state energy level (Fig. 39-5a) or by making two *separate* jumps by way of the $n = 2$ level (Figs. 39-5b and c).

The direct jump involves the same energy difference ΔE_{31} we found in (c). Then the wavelength is the same as we calculated in (c)—except now the wavelength is for light

that is emitted, not absorbed. Thus, the electron can jump directly to the ground state by emitting light of wavelength

$$\lambda = 4.12 \times 10^{-9} \text{ m.} \quad (\text{Answer})$$

Following the procedure of part (b), you can show that the energy differences for the jumps of Figs. 39-5b and c are

$$\Delta E_{32} = 3.016 \times 10^{-17} \text{ J} \quad \text{and} \quad \Delta E_{21} = 1.809 \times 10^{-17} \text{ J.}$$



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From Eq. 39-9, we then find that the wavelength of the light emitted in the first of these jumps (from $n = 3$ to $n = 2$) is

$$\lambda = 6.60 \times 10^{-9} \text{ m,} \quad (\text{Answer})$$

and the wavelength of the light emitted in the second of these jumps (from $n = 2$ to $n = 1$) is

$$\lambda = 1.10 \times 10^{-8} \text{ m.} \quad (\text{Answer})$$

39-4 Wave Functions of a Trapped Electron

If we solve Schrödinger's equation for an electron trapped in a one-dimensional infinite potential well of width L , we find that the wave functions for the electron are given by

$$\psi_n(x) = A \sin\left(\frac{n\pi}{L} x\right), \quad \text{for } n = 1, 2, 3, \dots, \quad (39-10)$$

for $0 \leq x \leq L$ (the wave function is zero outside that range). We shall soon evaluate the amplitude constant A in this equation.

Note that the wave functions $\psi_n(x)$ have the same form as the displacement functions $y_n(x)$ for a standing wave on a string stretched between rigid supports (see Eq. 39-2). We can picture an electron trapped in a one-dimensional well between infinite-potential walls as being a standing matter wave.

Probability of Detection

The wave function $\psi_n(x)$ cannot be detected or directly measured in any way—we cannot simply look inside the well to see the wave the way we can see, say, a wave in a bathtub of water. All we can do is insert a probe of some kind to try to detect the electron. At the instant of detection, the electron would materialize at the point of detection, at some position along the x axis within the well.

If we repeated this detection procedure at many positions throughout the well, we would find that the probability of detecting the electron is related to the probe's position x in the well. In fact, they are related by the *probability density* $\psi_n^2(x)$. Recall from Section 38-7 that in general the probability that a particle can be detected in a specified infinitesimal volume centered on a specified point is proportional to $|\psi_n|^2$. Here, with the electron trapped in a one-dimensional well, we are concerned only with detection of the electron along the x axis. Thus, the probability density $\psi_n^2(x)$ here is a probability per unit length along the x axis. (We can omit the absolute value sign here because $\psi_n(x)$ in Eq. 39-10 is a real quantity, not a complex one.) The probability $p(x)$ that an electron can be detected at position x within the well is

$$\begin{pmatrix} \text{probability } p(x) \\ \text{of detection in width } dx \\ \text{centered on position } x \end{pmatrix} = \begin{pmatrix} \text{probability density } \psi_n^2(x) \\ \text{at position } x \end{pmatrix} (\text{width } dx),$$

$$\text{or} \quad p(x) = \psi_n^2(x) dx. \quad (39-11)$$

From Eq. 39-10, we see that the probability density $\psi_n^2(x)$ is

$$\psi_n^2(x) = A^2 \sin^2\left(\frac{n\pi}{L} x\right), \quad \text{for } n = 1, 2, 3, \dots, \quad (39-12)$$

39-4 WAVE FUNCTIONS OF A TRAPPED ELECTRON

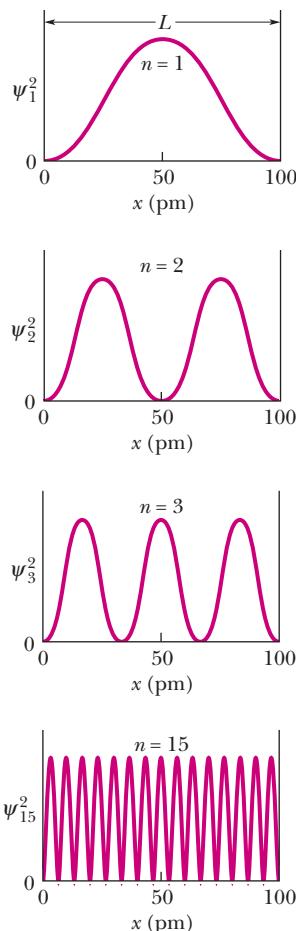
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for the range $0 \leq x \leq L$ (the probability density is zero outside that range). Figure 39-6 shows $\psi_n^2(x)$ for $n = 1, 2, 3$, and 15 for an electron in an infinite well whose width L is 100 pm.

To find the probability that the electron can be detected in any finite section of the well—say, between point x_1 and point x_2 —we must integrate $p(x)$ between those points. Thus, from Eqs. 39-11 and 39-12,

$$\begin{aligned} \left(\text{probability of detection} \right) &= \int_{x_1}^{x_2} p(x) dx \\ &= \int_{x_1}^{x_2} A^2 \sin^2\left(\frac{n\pi}{L}x\right) dx. \end{aligned} \quad (39-13)$$

The probability density must be zero at the infinite walls.



If classical physics prevailed, we would expect the trapped electron to be detectable with equal probabilities in all parts of the well. From Fig. 39-6 we see that it is not. For example, inspection of that figure or of Eq. 39-12 shows that for the state with $n = 2$, the electron is most likely to be detected near $x = 25$ pm and $x = 75$ pm. It can be detected with near-zero probability near $x = 0$, $x = 50$ pm, and $x = 100$ pm.

The case of $n = 15$ in Fig. 39-6 suggests that as n increases, the probability of detection becomes more and more uniform across the well. This result is an instance of a general principle called the **correspondence principle**:

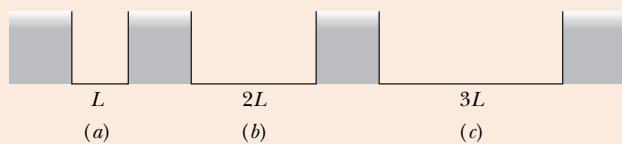


At large enough quantum numbers, the predictions of quantum physics merge smoothly with those of classical physics.

This principle, first advanced by Danish physicist Niels Bohr, holds for all quantum predictions.

**CHECKPOINT 2**

The figure shows three infinite potential wells of widths L , $2L$, and $3L$; each contains an electron in the state for



which $n = 10$. Rank the wells according to (a) the number of maxima for the probability density of the electron and (b) the energy of the electron, greatest first.

Normalization

The product $\psi_n^2(x) dx$ gives the probability that an electron in an infinite well can be detected in the interval of the x axis that lies between x and $x + dx$. We know that the electron must be *somewhere* in the infinite well; so it must be true that

$$\int_{-\infty}^{+\infty} \psi_n^2(x) dx = 1 \quad (\text{normalization equation}), \quad (39-14)$$

because the probability 1 corresponds to certainty. Although the integral is taken over the entire x axis, only the region from $x = 0$ to $x = L$ makes any contribution to the probability. Graphically, the integral in Eq. 39-14 represents the area under each of the plots of Fig. 39-6.

If we substitute $\psi_n^2(x)$ from Eq. 39-12 into Eq. 39-14, it is possible to assign a specific value to the amplitude constant A that appears in Eq. 39-12; namely, $A = \sqrt{2/L}$. This process of using Eq. 39-14 to evaluate the amplitude of a wave function is called **normalizing** the wave function. The process applies to *all* one-dimensional wave functions.

Fig. 39-6 The probability density $\psi_n^2(x)$ for four states of an electron trapped in a one-dimensional infinite well; their quantum numbers are $n = 1, 2, 3$, and 15. The electron is most likely to be found where $\psi_n^2(x)$ is greatest and least likely to be found where $\psi_n^2(x)$ is least.

Zero-Point Energy

Substituting $n = 1$ in Eq. 39-4 defines the state of lowest energy for an electron in an infinite potential well, the ground state. That is the state the confined electron will occupy unless energy is supplied to it to raise it to an excited state.

The question arises: Why can't we include $n = 0$ among the possibilities listed for n in Eq. 39-4? Putting $n = 0$ in this equation would indeed yield a ground-state energy of zero. However, putting $n = 0$ in Eq. 39-12 would also yield $\psi_n^2(x) = 0$ for all x , which we can interpret only to mean that there is no electron in the well. We know that there is; so $n = 0$ is not a possible quantum number.

It is an important conclusion of quantum physics that confined systems cannot exist in states with zero energy. They must always have a certain minimum energy called the **zero-point energy**.

We can make the zero-point energy as small as we like by making the infinite well wider—that is, by increasing L in Eq. 39-4 for $n = 1$. In the limit as $L \rightarrow \infty$, the zero-point energy E_1 approaches zero. In this limit, however, with an infinitely wide well, the electron is a free particle, no longer confined in the x direction. Also, because the energy of a free particle is not quantized, that energy can have any value, including zero. Only a confined particle must have a finite zero-point energy and can never be at rest.



CHECKPOINT 3

Each of the following particles is confined to an infinite well, and all four wells have the same width: (a) an electron, (b) a proton, (c) a deuteron, and (d) an alpha particle. Rank their zero-point energies, greatest first. The particles are listed in order of increasing mass.

Sample Problem

Detection probability in a 1D infinite potential well

A ground-state electron is trapped in the one-dimensional infinite potential well of Fig. 39-2, with width $L = 100 \text{ pm}$.

- (a) What is the probability that the electron can be detected in the left one-third of the well ($x_1 = 0$ to $x_2 = L/3$)?

KEY IDEAS

(1) If we probe the left one-third of the well, there is no guarantee that we will detect the electron. However, we can calculate the probability of detecting it with the integral of Eq. 39-13. (2) The probability very much depends on which state the electron is in—that is, the value of quantum number n .

Calculations: Because here the electron is in the ground state, we set $n = 1$ in Eq. 39-13. We also set the limits of integration as the positions $x_1 = 0$ and $x_2 = L/3$ and set the amplitude constant A as $\sqrt{2/L}$ (so that the wave function is normalized). We then see that

$$\left(\begin{array}{l} \text{probability of detection} \\ \text{in left one-third} \end{array} \right) = \int_0^{L/3} \frac{2}{L} \sin^2\left(\frac{1\pi}{L}x\right) dx.$$

We could find this probability by substituting $100 \times 10^{-12} \text{ m}$ for L and then using a graphing calculator or a computer math package to evaluate the integral. Here, however, we shall evaluate the integral “by hand.” First we switch to a new integration variable y :

$$y = \frac{\pi}{L}x \quad \text{and} \quad dx = \frac{L}{\pi} dy.$$

From the first of these equations, we find the new limits of integration to be $y_1 = 0$ for $x_1 = 0$ and $y_2 = \pi/3$ for $x_2 = L/3$. We then must evaluate

$$\text{probability} = \left(\frac{2}{L} \right) \left(\frac{L}{\pi} \right) \int_0^{\pi/3} (\sin^2 y) dy.$$

Using integral 11 in Appendix E, we then find

$$\text{probability} = \frac{2}{\pi} \left(\frac{y}{2} - \frac{\sin 2y}{4} \right)_0^{\pi/3} = 0.20.$$

Thus, we have

$$\left(\begin{array}{l} \text{probability of detection} \\ \text{in left one-third} \end{array} \right) = 0.20. \quad (\text{Answer})$$

That is, if we repeatedly probe the left one-third of the well, then on average we can detect the electron with 20% of the probes.

- (b) What is the probability that the electron can be detected in the middle one-third of the well?

Reasoning: We now know that the probability of detection in the left one-third of the well is 0.20. By symmetry, the probability of detection in the right one-third of the well is also 0.20. Because the electron is certainly in the well, the probability of detection in the entire well is 1. Thus, the probability of detection in the middle one-third of the well is

$$\left(\begin{array}{l} \text{probability of detection} \\ \text{in middle one-third} \end{array} \right) = 1 - 0.20 - 0.20 = 0.60. \quad (\text{Answer})$$



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Sample Problem**Normalizing wave functions in a 1D infinite potential well**

Evaluate the amplitude constant A in Eq. 39-10 for an infinite potential well extending from $x = 0$ to $x = L$.

KEY IDEA

The wave functions of Eq. 39-10 must satisfy the normalization requirement of Eq. 39-14, which states that the probability that the electron can be detected somewhere along the x axis is 1.

Calculations: Substituting Eq. 39-10 into Eq. 39-14 and taking the constant A outside the integral yield

$$A^2 \int_0^L \sin^2\left(\frac{n\pi}{L}x\right) dx = 1. \quad (39-15)$$

We have changed the limits of the integral from $-\infty$ and $+\infty$ to 0 and L because the wave function is zero outside these new limits (so there's no need to integrate out there).

We can simplify the indicated integration by changing the variable from x to the dimensionless variable y , where

$$y = \frac{n\pi}{L} x, \quad (39-16)$$

hence

$$dx = \frac{L}{n\pi} dy.$$



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When we change the variable, we must also change the integration limits (again). Equation 39-16 tells us that $y = 0$ when $x = 0$ and that $y = n\pi$ when $x = L$; thus 0 and $n\pi$ are our new limits. With all these substitutions, Eq. 39-15 becomes

$$A^2 \frac{L}{n\pi} \int_0^{n\pi} (\sin^2 y) dy = 1.$$

We can use integral 11 in Appendix E to evaluate the integral, obtaining the equation

$$\frac{A^2 L}{n\pi} \left[\frac{y}{2} - \frac{\sin 2y}{4} \right]_0^{n\pi} = 1.$$

Evaluating at the limits yields

$$\frac{A^2 L}{n\pi} \frac{n\pi}{2} = 1;$$

thus $A = \sqrt{\frac{2}{L}}$. (Answer) (39-17)

This result tells us that the dimension for A^2 , and thus for $\psi_n^2(x)$, is an inverse length. This is appropriate because the probability density of Eq. 39-12 is a probability *per unit length*.

39-5 An Electron in a Finite Well

A potential energy well of infinite depth is an idealization. Figure 39-7 shows a realizable potential energy well—one in which the potential energy of an electron outside the well is not infinitely great but has a finite positive value U_0 , called the **well depth**. The analogy between waves on a stretched string and matter waves fails us for wells of finite depth because we can no longer be sure that matter wave nodes exist at $x = 0$ and at $x = L$. (As we shall see, they don't.)

To find the wave functions describing the quantum states of an electron in the finite well of Fig. 39-7, we *must* resort to Schrödinger's equation, the basic equation of quantum physics. From Section 38-7 recall that, for motion in one dimension, we use Schrödinger's equation in the form of Eq. 38-15:

$$\frac{d^2\psi}{dx^2} + \frac{8\pi^2 m}{h^2} [E - U(x)]\psi = 0. \quad (39-18)$$

Rather than attempting to solve this equation for the finite well, we simply state the results for particular numerical values of U_0 and L . Figure 39-8 shows three

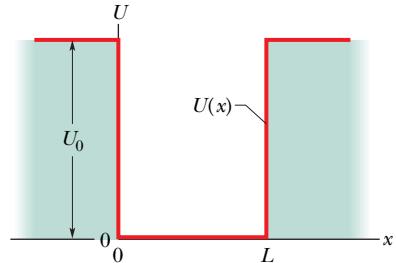
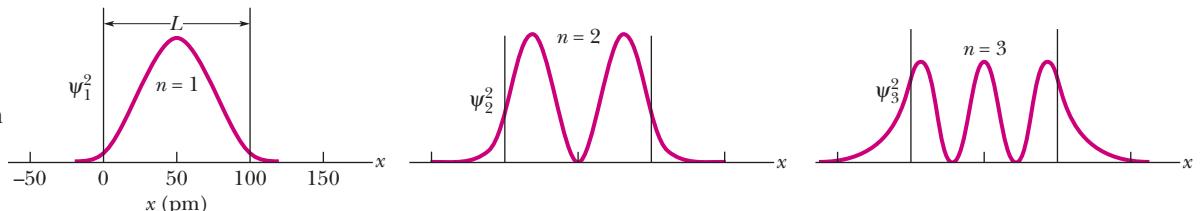


Fig. 39-7 A finite potential energy well. The depth of the well is U_0 and its width is L . As in the infinite potential well of Fig. 39-2, the motion of the trapped electron is restricted to the x direction.

Fig. 39-8 The first three probability densities $\psi_n^2(x)$ for an electron confined to a finite potential well of depth $U_0 = 450$ eV and width $L = 100$ pm. Only states $n = 1, 2, 3$, and 4 are allowed.



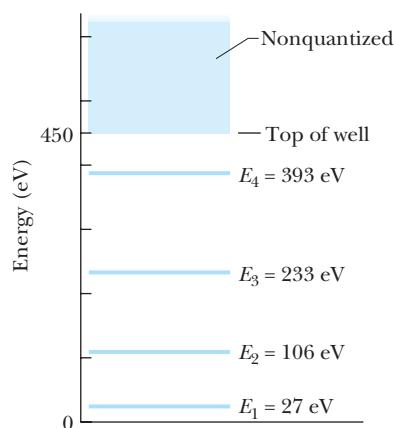


Fig. 39-9 The energy-level diagram corresponding to the probability densities of Fig. 39-8. If an electron is trapped in the finite potential well, it can have only the energies corresponding to $n = 1, 2, 3$, and 4 . If it has an energy of 450 eV or greater, it is not trapped and its energy is not quantized.

results as graphs of $\psi_n^2(x)$, the probability density, for a well with $U_0 = 450$ eV and $L = 100$ pm.

The probability density $\psi_n^2(x)$ for each graph in Fig. 39-8 satisfies Eq. 39-14, the normalization equation; so we know that the areas under all three probability density plots are numerically equal to 1.

If you compare Fig. 39-8 for a finite well with Fig. 39-6 for an infinite well, you will see one striking difference: For a finite well, the electron matter wave penetrates the walls of the well—into a region in which Newtonian mechanics says the electron cannot exist. This penetration should not be surprising because we saw in Section 38-9 that an electron can tunnel through a potential energy barrier. “Leaking” into the walls of a finite potential energy well is a similar phenomenon. From the plots of ψ^2 in Fig. 39-8, we see that the leakage is greater for greater values of quantum number n .

Because a matter wave *does* leak into the walls of a finite well, the wavelength λ for any given quantum state is greater when the electron is trapped in a finite well than when it is trapped in an infinite well. Equation 39-3 ($\lambda = h/\sqrt{2mE}$) then tells us that the energy E for an electron in any given state is less in the finite well than in the infinite well.

That fact allows us to approximate the energy-level diagram for an electron trapped in a finite well. As an example, we can approximate the diagram for the finite well of Fig. 39-8, which has width $L = 100$ pm and depth $U_0 = 450$ eV. The energy-level diagram for an *infinite* well of that width is shown in Fig. 39-3. First we remove the portion of Fig. 39-3 above 450 eV. Then we shift the remaining four energy levels down, shifting the level for $n = 4$ the most because the wave leakage into the walls is greatest for $n = 4$. The result is approximately the energy-level diagram for the finite well. The actual diagram is Fig. 39-9.

In that figure, an electron with an energy greater than U_0 (= 450 eV) has too much energy to be trapped in the finite well. Thus, it is not confined, and its energy is not quantized; that is, its energy is not restricted to certain values. To reach this *nonquantized* portion of the energy-level diagram and thus to be free, a trapped electron must somehow obtain enough energy to have a mechanical energy of 450 eV or greater.

Sample Problem

Electron escaping from a finite potential well

Suppose a finite well with $U_0 = 450$ eV and $L = 100$ pm confines a single electron in its ground state.

- (a) What wavelength of light is needed to barely free the electron from the potential well if the electron absorbs a single photon from the light?

KEY IDEA

For the electron to escape from the potential well, it must receive enough energy to put it into the nonquantized energy region of Fig. 39-9. Thus, it must end up with an energy of at least U_0 (= 450 eV).

Barely escaping: The electron is initially in its ground state, with an energy of $E_1 = 27$ eV. So, to barely become free, it must receive an energy of

$$U_0 - E_1 = 450 \text{ eV} - 27 \text{ eV} = 423 \text{ eV}.$$

Thus the photon must have this much energy. From Eq.

39-6 ($hf = E_{\text{high}} - E_{\text{low}}$), with c/λ substituted for f , we write

$$\frac{hc}{\lambda} = U_0 - E_1,$$

from which we find

$$\begin{aligned}\lambda &= \frac{hc}{U_0 - E_1} \\ &= \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s})}{(423 \text{ eV})(1.60 \times 10^{-19} \text{ J/eV})} \\ &= 2.94 \times 10^{-9} \text{ m} = 2.94 \text{ nm}. \quad (\text{Answer})\end{aligned}$$

Thus, if $\lambda = 2.94$ nm, the electron just barely escapes.

- (b) Can the ground-state electron absorb light with $\lambda = 2.00$ nm? If so, what then is the electron's energy?

KEY IDEAS

1. In (a) we found that light of 2.94 nm will just barely free the electron from the potential well.

2. We are now considering light with a shorter wavelength of 2.00 nm and thus a greater energy per photon ($hf = hc/\lambda$).
3. Hence, the electron *can* absorb a photon of this light. The energy transfer will not only free the electron but will also provide it with more kinetic energy. Further, because the electron is then no longer trapped, its energy is not quantized.

More than escaping: The energy transferred to the electron is the photon energy:

$$hf = h \frac{c}{\lambda} = \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s})}{2.00 \times 10^{-9} \text{ m}} \\ = 9.95 \times 10^{-17} \text{ J} = 622 \text{ eV.}$$

From (a), the energy required to just barely free the electron from the potential well is $U_0 - E_1$ (= 423 eV). The remainder of the 622 eV goes to kinetic energy. Thus, the kinetic energy of the freed electron is

$$K = hf - (U_0 - E_1) \\ = 622 \text{ eV} - 423 \text{ eV} = 199 \text{ eV. (Answer)}$$



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39-6 More Electron Traps

Here we discuss three types of artificial electron traps.

Nanocrystallites

Perhaps the most direct way to construct a potential energy well in the laboratory is to prepare a sample of a semiconducting material in the form of a powder whose granules are small—in the nanometer range—and of uniform size. Each such granule—each **nanocrystallite**—acts as a potential well for the electrons trapped within it.

Equation 39-4 ($E = (h^2/8mL^2)n^2$) shows that we can increase the energy-level values of an electron trapped in an infinite well by reducing the width L of the well. This would also shift the photon energies that the well can absorb to higher values and thus shift the corresponding wavelengths to shorter values.

These general results are also true for a well formed by a nanocrystallite. A given nanocrystallite can absorb photons with an energy above a certain threshold energy E_t ($= hf_t$) and thus wavelengths below a corresponding threshold wavelength

$$\lambda_t = \frac{c}{f_t} = \frac{ch}{E_t}.$$

Light with any wavelength longer than λ_t is scattered by the nanocrystallite instead of being absorbed. The color we attribute to the nanocrystallite is then determined by the wavelength composition of the scattered light we intercept.

If we reduce the size of the nanocrystallite, the value of E_t is increased, the value of λ_t is decreased, and the light that is scattered to us changes in its wavelength composition. Thus, the color we attribute to the nanocrystallite changes. As an example, Fig. 39-10 shows two samples of the semiconductor cadmium selenide, each consisting of a powder of nanocrystallites of uniform size. The lower sample scatters light at the red end of the spectrum. The upper sample differs from the lower sample *only* in that the upper sample is composed of smaller nanocrystallites. For this reason its threshold energy E_t is greater and, from above, its threshold wavelength λ_t is shorter, in the green range of visible light. Thus, the sample now scatters both red and yellow. Because the yellow component happens to be brighter, the sample's color is now dominated by the yellow. The striking contrast in color between the two samples is compelling evidence of the quantization of the energies of trapped electrons and the dependence of these energies on the size of the electron trap.



Fig. 39-10 Two samples of powdered cadmium selenide, a semiconductor, differing only in the size of their granules. Each granule serves as an electron trap. The lower sample has the larger granules and consequently the smaller spacing between energy levels and the lower photon energy threshold for the absorption of light. Light not absorbed is scattered, causing the sample to scatter light of greater wavelength and appear red. The upper sample, because of its smaller granules, and consequently its larger level spacing and its larger energy threshold for absorption, appears yellow. (From Scientific American, January 1993, page 122. Reproduced with permission of Michael Steigerwald, Bell Labs–Lucent Technologies)

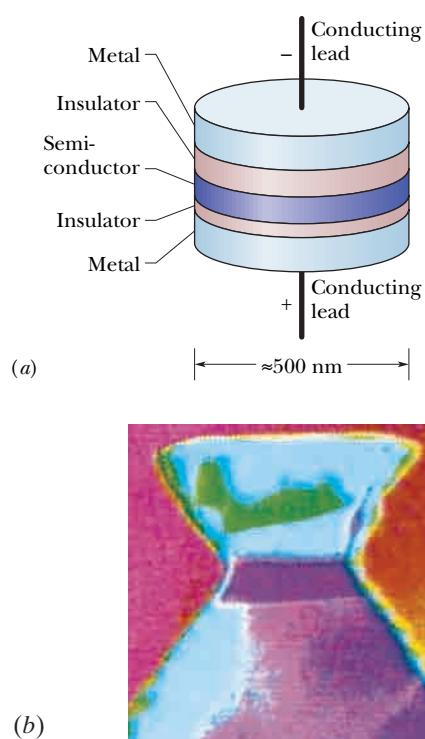


Fig. 39-11 A quantum dot, or “artificial atom.” (a) A central semiconducting layer forms a potential energy well in which electrons are trapped. The lower insulating layer is thin enough to allow electrons to be added to or removed from the central layer by barrier tunneling if an appropriate voltage is applied between the leads. (b) A photograph of an actual quantum dot. The central purple band is the electron confinement region.
(From Scientific American, September 1995, page 67. Image reproduced with permission of H. Temkin, Texas Tech University)

Quantum Dots

The highly developed techniques used to fabricate computer chips can be used to construct, atom by atom, individual potential energy wells that behave, in many respects, like artificial atoms. These **quantum dots**, as they are usually called, have promising applications in electron optics and computer technology.

In one such arrangement, a “sandwich” is fabricated in which a thin layer of a semiconducting material, shown in purple in Fig. 39-11a, is deposited between two insulating layers, one of which is much thinner than the other. Metal end caps with conducting leads are added at both ends. The materials are chosen to ensure that the potential energy of an electron in the central layer is less than it is in the two insulating layers, causing the central layer to act as a potential energy well. Figure 39-11b is a photograph of an actual quantum dot; the well in which individual electrons can be trapped is the purple region.

The lower (but not the upper) insulating layer in Fig. 39-11a is thin enough to permit electrons to tunnel through it if an appropriate potential difference is applied between the leads. In this way the number of electrons confined to the well can be controlled. The arrangement does indeed behave like an artificial atom with the property that the number of electrons it contains can be controlled. Quantum dots can be constructed in two-dimensional arrays that could well form the basis for computing systems of great speed and storage capacity.

Quantum Corrals

When a scanning tunneling microscope (described in Section 38-9) is in operation, its tip exerts a small force on isolated atoms that may be located on an otherwise smooth surface. By careful manipulation of the position of the tip, such isolated atoms can be “dragged” across the surface and deposited at another location. Using this technique, scientists at IBM’s Almaden Research Center moved iron atoms across a carefully prepared copper surface, forming the atoms into a circle (Fig. 39-12), which they named a **quantum corral**. Each iron atom in the circle is nestled in a hollow in the copper surface, equidistant from three nearest-neighbor copper atoms. The corral was fabricated at a low temperature (about 4 K) to minimize the tendency of the iron atoms to move randomly about on the surface because of their thermal energy.

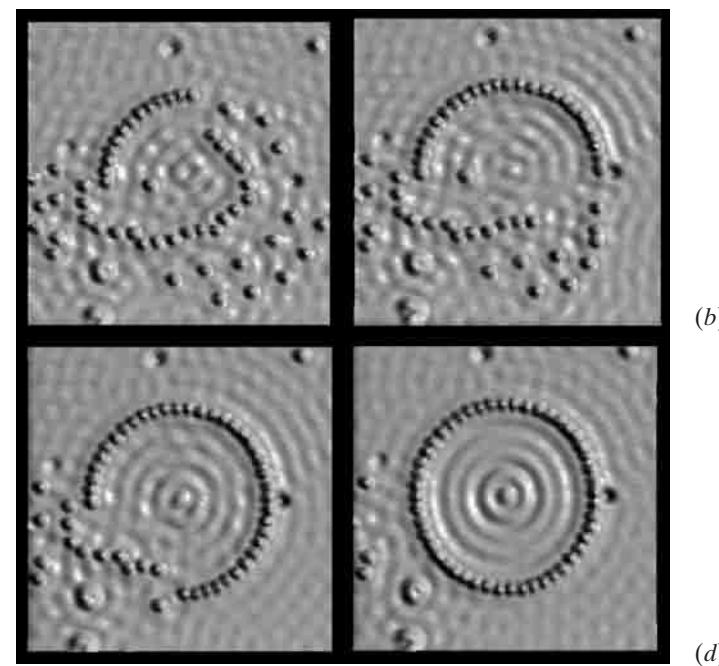


Fig. 39-12 A quantum corral during four stages of construction. Note the appearance of ripples caused by electrons trapped in the corral when it is almost complete. (Courtesy of International Business Machines Corporation, Almaden Research Center)

39-7 TWO- AND THREE-DIMENSIONAL ELECTRON TRAPS

1095

The ripples within the corral are due to matter waves associated with electrons that can move over the copper surface but are largely trapped in the potential well of the corral. The dimensions of the ripples are in excellent agreement with the predictions of quantum theory.

39-7 Two- and Three-Dimensional Electron Traps

In the next section, we shall discuss the hydrogen atom as being a three-dimensional finite potential well. As a warm-up for the hydrogen atom, let us extend our discussion of infinite potential wells to two and three dimensions.

Rectangular Corral

Figure 39-13 shows the rectangular area to which an electron can be confined by the two-dimensional version of Fig. 39-2—a two-dimensional infinite potential well of widths L_x and L_y that forms a rectangular corral. The corral might be on the surface of a body that somehow prevents the electron from moving parallel to the z axis and thus from leaving the surface. You have to imagine infinite potential energy functions (like $U(x)$ in Fig. 39-2) along each side of the corral, keeping the electron within the corral.

Solution of Schrödinger's equation for the rectangular corral of Fig. 39-13 shows that, for the electron to be trapped, its matter wave must fit into each of the two widths separately, just as the matter wave of a trapped electron must fit into a one-dimensional infinite well. This means the wave is separately quantized in width L_x and in width L_y . Let n_x be the quantum number for which the matter wave fits into width L_x , and let n_y be the quantum number for which the matter wave fits into width L_y . As with a one-dimensional potential well, these quantum numbers can be only positive integers. We can extend Eqs. 39-10 and 39-17 to write the normalized wave function as

$$\psi_{nx,ny} = \sqrt{\frac{2}{L_x}} \sin\left(\frac{n_x\pi}{L} x\right) \sqrt{\frac{2}{L_y}} \sin\left(\frac{n_y\pi}{L} y\right), \quad (39-19)$$

The energy of the electron depends on both quantum numbers and is the sum of the energy the electron would have if it were confined along the x axis alone and the energy it would have if it were confined along the y axis alone. From Eq. 39-4, we can write this sum as

$$E_{nx,ny} = \left(\frac{h^2}{8mL_x^2}\right)n_x^2 + \left(\frac{h^2}{8mL_y^2}\right)n_y^2 = \frac{h^2}{8m} \left(\frac{n_x^2}{L_x^2} + \frac{n_y^2}{L_y^2}\right). \quad (39-20)$$

Excitation of the electron by photon absorption and de-excitation of the electron by photon emission have the same requirements as for one-dimensional traps. Now, however, two quantum numbers (n_x and n_y) are involved. Because of that, different states might have the same energy; such states and their energy levels are said to be *degenerate*.

Rectangular Box

An electron can also be trapped in a three-dimensional infinite potential well—a *box*. If the box is rectangular as in Fig. 39-14, then Schrödinger's equation shows us that we can write the energy of the electron as

$$E_{nx,ny,nz} = \frac{h^2}{8m} \left(\frac{n_x^2}{L_x^2} + \frac{n_y^2}{L_y^2} + \frac{n_z^2}{L_z^2}\right). \quad (39-21)$$

Here n_z is a third quantum number, for fitting the matter wave into width L_z .

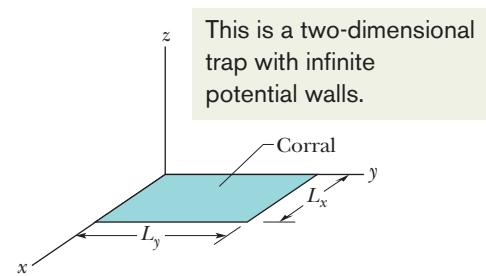


Fig. 39-13 A rectangular corral—a two-dimensional version of the infinite potential well of Fig. 39-2—with widths L_x and L_y .

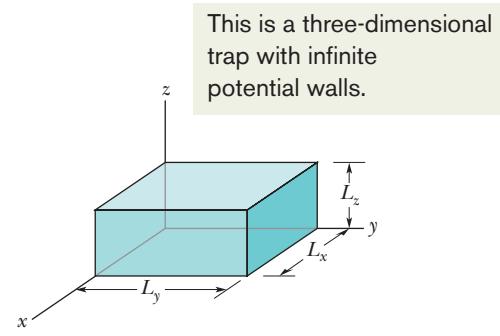


Fig. 39-14 A rectangular box—a three-dimensional version of the infinite potential well of Fig. 39-2—with widths L_x , L_y , and L_z .



CHECKPOINT 4

In the notation of Eq. 39-20, is $E_{0,0}$, $E_{1,0}$, $E_{0,1}$, or $E_{1,1}$ the ground-state energy of an electron in a (two-dimensional) rectangular corral?

Sample Problem**Energy levels in a 2D infinite potential well**

An electron is trapped in a square corral that is a two-dimensional infinite potential well (Fig. 39-13) with widths $L_x = L_y$.

- (a) Find the energies of the lowest five possible energy levels for this trapped electron, and construct the corresponding energy-level diagram.

KEY IDEA

Because the electron is trapped in a two-dimensional well that is rectangular, the electron's energy depends on two quantum numbers, n_x and n_y , according to Eq. 39-20.

Energy levels: Because the well here is square, we can let the widths be $L_x = L_y = L$. Then Eq. 39-20 simplifies to

$$E_{nx,ny} = \frac{h^2}{8mL^2} (n_x^2 + n_y^2). \quad (39-22)$$

The lowest energy states correspond to low values of the quantum numbers n_x and n_y , which are the positive integers $1, 2, \dots, \infty$. Substituting those integers for n_x and n_y in Eq. 39-22, starting with the lowest value 1, we can obtain the energy values as listed in Table 39-1. There we can see that

Table 39-1**Energy Levels**

n_x	n_y	Energy ^a	n_x	n_y	Energy ^a
1	3	10	2	4	20
3	1	10	4	2	20
2	2	8	3	3	18
1	2	5	1	4	17
2	1	5	4	1	17
1	1	2	2	3	13
			3	2	13

^aIn multiples of $h^2/8mL^2$.

These are the lowest five energy levels allowed the electron. Different quantum states may have the same energy.

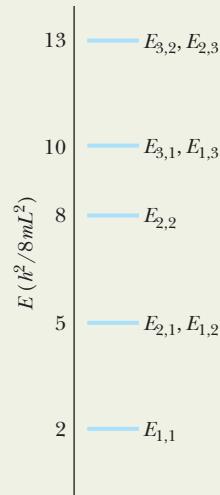


Fig. 39-15 Energy-level diagram for an electron trapped in a square corral.

several of the pairs of quantum numbers (n_x, n_y) give the same energy. For example, the (1, 2) and (2, 1) states both have an energy of $5(h^2/8mL^2)$. Each such pair is associated with degenerate energy levels. Note also that, perhaps surprisingly, the (4, 1) and (1, 4) states have less energy than the (3, 3) state.

From Table 39-1 (carefully keeping track of degenerate levels), we can construct the energy-level diagram of Fig. 39-15.

- (b) As a multiple of $h^2/8mL^2$, what is the energy difference between the ground state and the third excited state?

Energy difference: From Fig. 39-15, we see that the ground state is the (1, 1) state, with an energy of $2(h^2/8mL^2)$. We also see that the third excited state (the third state up from the ground state in the energy-level diagram) is the degenerate (1, 3) and (3, 1) states, with an energy of $10(h^2/8mL^2)$. Thus, the difference ΔE between these two states is

$$\Delta E = 10\left(\frac{h^2}{8mL^2}\right) - 2\left(\frac{h^2}{8mL^2}\right) = 8\left(\frac{h^2}{8mL^2}\right).$$

(Answer)



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39-8 The Bohr Model of the Hydrogen Atom

We now move from artificial and fictitious electron traps to natural ones—atoms. In this chapter we focus on the simplest example, a hydrogen atom, which contains an electron that is trapped to be near the proton, which forms the atom's central nucleus. Here we do not consider anything about the nucleus. Rather, we simply use the fact that the negatively charged electron is attracted by the Coulomb force to the positively charged proton. Because the proton mass is much greater than the electron mass, we assume that the proton is fixed in place.

39-8 THE BOHR MODEL OF THE HYDROGEN ATOM

1097

We have now discussed at length that confinement of an electron means that the electron's energy E is quantized and thus so is any change ΔE in its energy. In this section we want to calculate the quantized energies of the electron in a hydrogen atom. Before we apply the wave approach we used in infinite and finite potential wells, however, let's explore the hydrogen atom at the dawn of quantum physics, when physicists first discovered that atoms are quantized systems.

By the early 1900s, scientists understood that matter came in tiny pieces called atoms and that an atom of hydrogen contained positive charge $+e$ at its center and negative charge $-e$ (an electron) outside that center. However, no one understood why the electrical attraction between the electron and the positive charge did not simply cause the two to collapse together.

One clue came from the fact that a hydrogen atom cannot emit and absorb all wavelengths of visible light. Rather, it can emit and absorb only four particular wavelengths from the visible range. By guesswork, Johann Balmer devised a formula that gave those wavelengths:

$$\frac{1}{\lambda} = R \left(\frac{1}{2^2} - \frac{1}{n^2} \right), \quad \text{for } n = 3, 4, 5, \text{ and } 6, \quad (39-23)$$

where R is a constant. However, no one knew why this formula gave the right wavelengths or why no other visible wavelengths are emitted and absorbed.

No one knew until 1913, when Bohr saw Balmer's equation and quickly realized that he could derive it if he made several bold (completely unjustified) assumptions: (1) The electron in a hydrogen atom orbits the nucleus in a circle much like Earth orbits the Sun (Fig. 39-16a). (2) The magnitude of the angular momentum \vec{L} of the electron in its orbit is restricted to the values

$$L = n\hbar, \quad \text{for } n = 1, 2, 3, \dots, \quad (39-24)$$

where \hbar (h-bar) is $h/2\pi$ and n is a quantum number. Let's see the results.

The Orbital Radius Is Quantized in the Bohr Model

Let's examine the orbital motion of the electron in the Bohr model. The force holding the electron in an orbit of radius r is the Coulomb force. From Eq. 21-1, we know that the magnitude of this force is

$$F = k \frac{|q_1||q_2|}{r^2},$$

with $k = 1/4\pi\epsilon_0$. Here q_1 is the charge $-e$ of the electron and q_2 is the charge $+e$ of the nucleus (the proton). The electron's acceleration is the centripetal acceleration, with a magnitude given by $a = v^2/r$, where v is the electron's speed. Both force \vec{F} and acceleration \vec{a} are radially inward (the negative direction on a radial axis), toward the nucleus (Fig. 39-16b). Thus, we can write Newton's second law ($F = ma$) for a radial axis as

$$-\frac{1}{4\pi\epsilon_0} \frac{e^2}{r^2} = m \left(-\frac{v^2}{r} \right), \quad (39-25)$$

where m is the electron mass.

We next introduce quantization by using Bohr's assumption expressed in Eq. 39-24. From Eq. 11-19, the magnitude ℓ of the angular momentum of a particle of mass m and speed v moving in a circle of radius r is $\ell = rmv \sin \phi$, where ϕ (the angle between \vec{r} and \vec{v}) is 90° . Replacing L in Eq. 39-24 with $rmv \sin 90^\circ$ gives us

$$rmv = n\hbar,$$

or

$$v = \frac{n\hbar}{rm}. \quad (39-26)$$

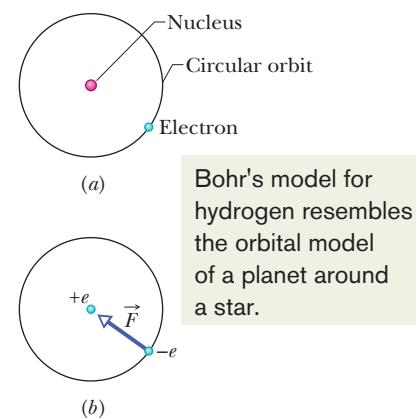


Fig. 39-16 (a) Circular orbit of an electron in the Bohr model of the hydrogen atom. (b) The Coulomb force \vec{F} on the electron is directed radially inward toward the nucleus.

Substituting this equation into Eq. 39-25, replacing \hbar with $h/2\pi$, and rearranging, we find

$$r = \frac{h^2 \epsilon_0}{\pi m e^2} n^2, \quad \text{for } n = 1, 2, 3, \dots \quad (39-27)$$

We can rewrite this as

$$r = an^2, \quad \text{for } n = 1, 2, 3, \dots, \quad (39-28)$$

$$\text{where } a = \frac{h^2 \epsilon_0}{\pi m e^2} = 5.291\ 772 \times 10^{-11} \text{ m} \approx 52.92 \text{ pm}. \quad (39-29)$$

These last three equations tell us that, in the *Bohr model of the hydrogen atom*, the electron's orbital radius r is quantized and the smallest possible orbital radius (for $n = 1$) is a , which is now called the *Bohr radius*. According to the Bohr model, the electron cannot get any closer to the nucleus than orbital radius a , and that is why the attraction between electron and nucleus does not simply collapse them together.

Orbital Energy Is Quantized

Let's next find the energy of the hydrogen atom according to the Bohr model. The electron has kinetic energy $K = \frac{1}{2}mv^2$, and the electron–nucleus system has electric potential energy $U = q_1q_2/4\pi\epsilon_0 r$ (Eq. 24-43). Again, let q_1 be the electron's charge $-e$ and q_2 be the nuclear charge $+e$. Then the mechanical energy is

$$\begin{aligned} E &= K + U \\ &= \frac{1}{2}mv^2 + \left(-\frac{1}{4\pi\epsilon_0} \frac{e^2}{r} \right). \end{aligned} \quad (39-30)$$

Solving Eq. 39-25 for mv^2 and substituting the result in Eq. 39-30 lead to

$$E = -\frac{1}{8\pi\epsilon_0} \frac{e^2}{r}. \quad (39-31)$$

Next, replacing r with its equivalent from Eq. 39-27, we have

$$E_n = -\frac{me^4}{8\epsilon_0^2 h^2} \frac{1}{n^2}, \quad \text{for } n = 1, 2, 3, \dots, \quad (39-32)$$

where the subscript n on E signals that we have now quantized the energy. Evaluating the constants in Eq. 39-32 gives us

$$E_n = -\frac{2.180 \times 10^{-18} \text{ J}}{n^2} = -\frac{13.61 \text{ eV}}{n^2}, \quad \text{for } n = 1, 2, 3, \dots. \quad (39-33)$$

This equation tells us that the energy E_n of the hydrogen atom is quantized; that is, E_n is restricted by its dependence on the quantum number n . Because the nucleus is assumed to be fixed in place and only the electron has motion, we can assign the energy values of Eq. 39-33 either to the atom as a whole or to the electron alone.

Energy Changes

The energy of a hydrogen atom (or, equivalently, of its electron) changes when the atom emits or absorbs light. As we have seen several times since Eq. 39-6, emission and absorption involve a quantum of light according to

$$hf = \Delta E = E_{\text{high}} - E_{\text{low}}. \quad (39-34)$$

Let's make three changes to Eq. 39-34. On the left side, we substitute c/λ for f . On the right side, we use Eq. 39-32 twice to replace the energy terms. Then,

39-9 SCHRÖDINGER'S EQUATION AND THE HYDROGEN ATOM

1099

with a simple rearrangement, we have

$$\frac{1}{\lambda} = -\frac{me^4}{8\varepsilon_0^2 h^3 c} \left(\frac{1}{n_{\text{high}}^2} - \frac{1}{n_{\text{low}}^2} \right). \quad (39-35)$$

We can rewrite this as

$$\frac{1}{\lambda} = R \left(\frac{1}{n_{\text{low}}^2} - \frac{1}{n_{\text{high}}^2} \right), \quad (39-36)$$

in which

$$R = \frac{me^4}{8\varepsilon_0^2 h^3 c} = 1.097\,373 \times 10^7 \text{ m}^{-1} \quad (39-37)$$

is now known as the *Rydberg constant*.

Compare Eq. 39-36 from the Bohr model with Eq. 39-23 from Balmer's work. In Eq. 39-36, if we replace n_{low} with 2 and then restrict n_{high} to be 3, 4, 5, and 6, we have Balmer's equation. This match was a triumph for Bohr and ushered in the quantum physics of atoms. The triumph was short-lived, however, because even though the Bohr model gives the correct emission and absorption wavelengths for the hydrogen atom, the model is *not* correct because the electron does *not* orbit the nucleus like a planet orbiting the Sun. Indeed, researchers found little success in extending the Bohr model to atoms more complicated than hydrogen. The reason for this lack of success is that an electron trapped in any atom is a matter wave confined to a potential well, and to find the resulting quantized energy values we must apply Schrödinger's equation to the electron.

39-9 Schrödinger's Equation and the Hydrogen Atom

The potential well of a hydrogen atom depends on the electrical potential energy function

$$U(r) = \frac{-e^2}{4\pi\varepsilon_0 r}. \quad (39-38)$$

Because this well is three-dimensional, it is more complex than our previous one- and two-dimensional wells. Because this well is finite, it is more complex than the three-dimensional well of Fig. 39-14. Moreover, it does not have sharply defined walls. Rather, its walls vary in depth with radial distance r . Figure 39-17 is probably the best we can do in drawing the hydrogen potential well, but even that drawing takes much effort to interpret.

Energy Levels and Spectra of the Hydrogen Atom

Although we shall not do so here, we can apply Schrödinger's equation for an electron trapped in the potential well given by Eq.

39-38. In doing so, we would find that the energy values are quantized and that, amazingly, those values are given by Eq. 39-33 just as for the (incorrect) Bohr model. Thus, changes ΔE in energy due to emission or absorption of light are given by Eq. 39-34, and the wavelengths corresponding to ΔE are given by Eq. 39-36. Let's explore these results.

Figure 39-18a shows the energy levels corresponding to various values of n in Eq. 39-33. The lowest level, for $n = 1$, is the ground state of hydrogen. Higher levels correspond to excited states, just as we saw for our simpler potential traps. Note several differences, however. (1) The energy levels now have negative values rather than the positive values we previously chose in, for instance, Figs. 39-3 and 39-9. (2) The levels now become progressively closer as we move to

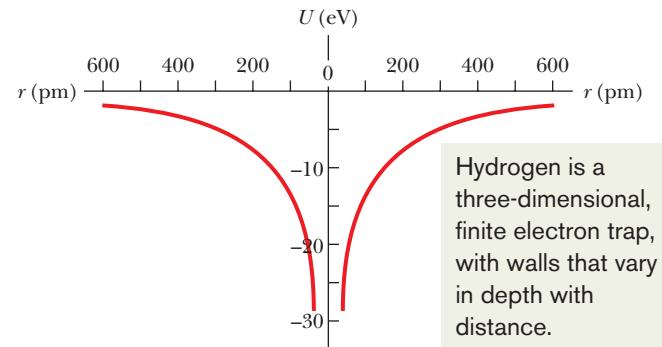


Fig. 39-17 The potential energy U of a hydrogen atom as a function of the separation r between the electron and the central proton. The plot is shown twice (on the left and on the right) to suggest the three-dimensional spherically symmetric trap in which the electron is confined.

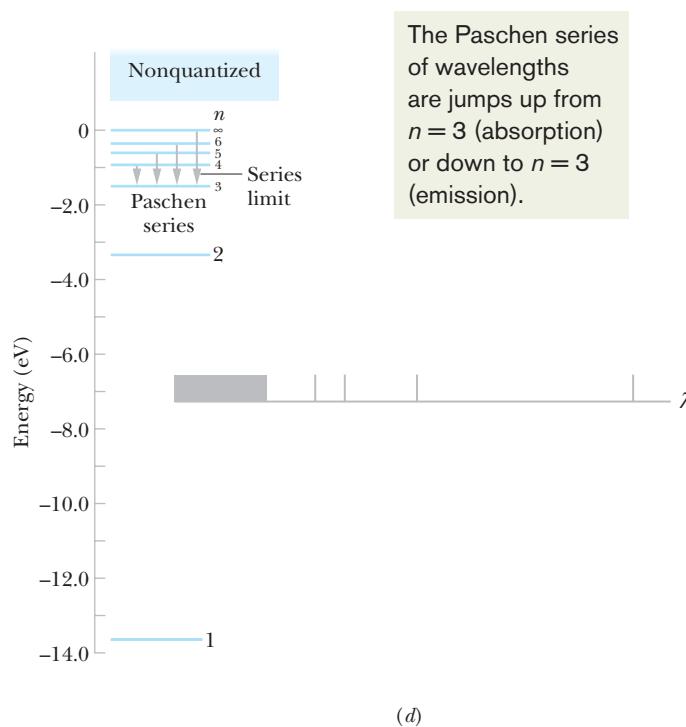
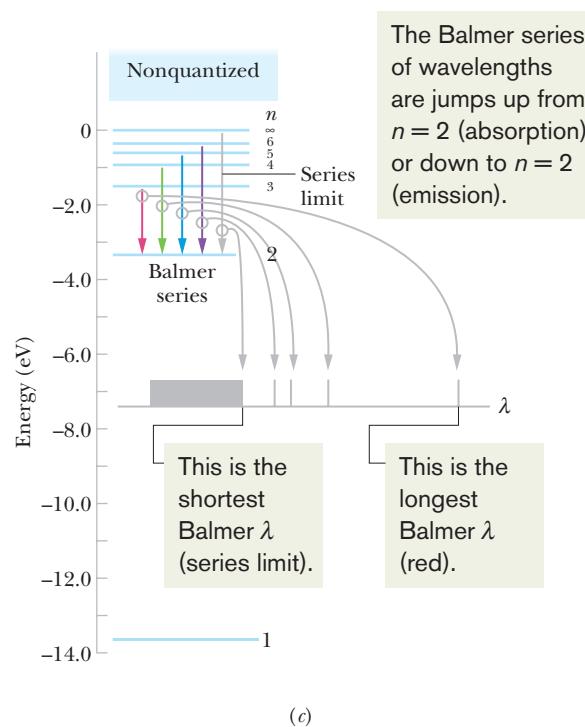
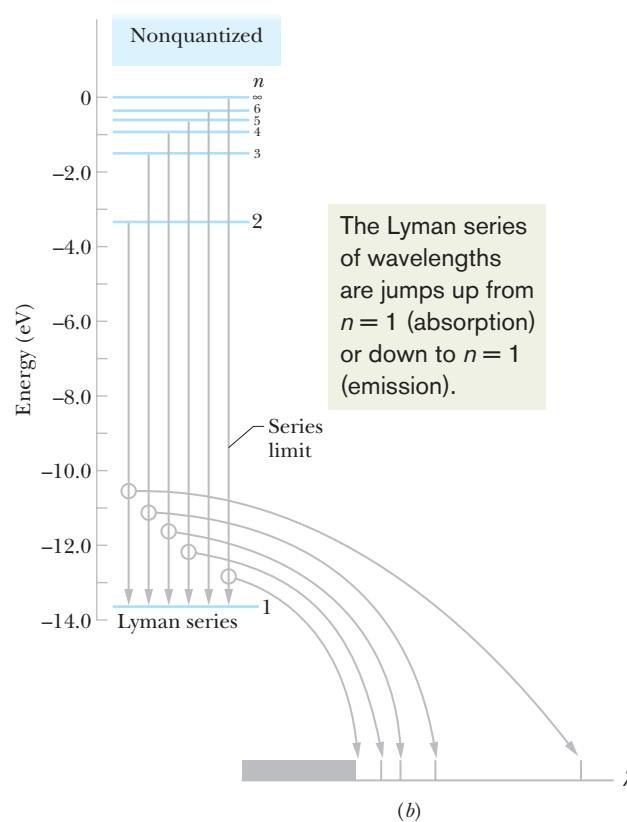
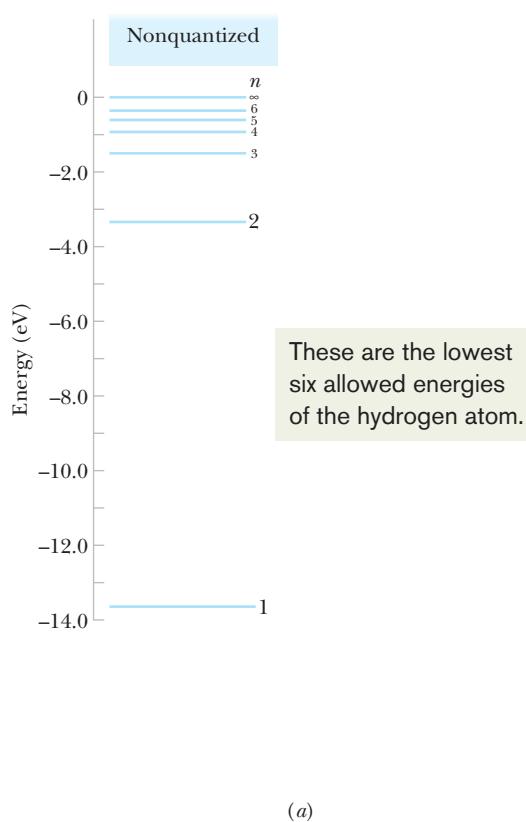


Fig. 39-18 (a) An energy-level diagram for the hydrogen atom. Some of the transitions for (b) the Lyman series, (c) the Balmer series, and (d) the Paschen series. For each, the longest four wavelengths and the series-limit wavelength are plotted on a wavelength axis. Any wavelength shorter than the series-limit wavelength is allowed.

39-9 SCHRÖDINGER'S EQUATION AND THE HYDROGEN ATOM

1101

higher levels. (3) The energy for the greatest value of n —namely, $n = \infty$ —is now $E_\infty = 0$. For any energy greater than $E_\infty = 0$, the electron and proton are not bound together (there is no hydrogen atom), and the $E > 0$ region in Fig. 39-18a is like the nonquantized region for the finite well of Fig. 39-9.

A hydrogen atom can jump between quantized energy levels by emitting or absorbing light at the wavelengths given by Eq. 39-36. Any such wavelength is often called a *line* because of the way it is detected with a spectroscope; thus, a hydrogen atom has *absorption lines* and *emission lines*. A collection of such lines, such as in those in the visible range, is called a **spectrum** of the hydrogen atom.

The lines for hydrogen are said to be grouped into *series*, according to the level at which upward jumps start and downward jumps end. For example, the emission and absorption lines for all possible jumps up from the $n = 1$ level and down to the $n = 1$ level are said to be in the *Lyman series* (Fig. 39-18b), named after the person who first studied those lines. Further, we can say that the Lyman series has a *home-base level* of $n = 1$. Similarly, the *Balmer series* has a home-base level of $n = 2$ (Fig. 39-18c), and the *Paschen series* has a home-base level of $n = 3$ (Fig. 39-18d).

Some of the downward quantum jumps for these three series are shown in Fig. 39-18. Four lines in the Balmer series are in the visible range and are the ones Balmer studied. They are represented in Fig. 39-18c with arrows corresponding to their colors. The shortest of those arrows represents the shortest jump in the series, from the $n = 3$ level to the $n = 2$ level. Thus, that jump involves the smallest change in the electron's energy and the smallest amount of emitted photon energy for the series. The emitted light is red. The next jump in the series, from $n = 4$ to $n = 2$, is longer, the photon energy is greater, the wavelength of the emitted light is shorter, and the light is green. The third, fourth, and fifth arrows represent longer jumps and shorter wavelengths. For the fifth jump, the emitted light is in the ultraviolet range and thus is not visible.

The *series limit* of a series is the line produced by the jump between the home-base level and the highest energy level, which is the level with quantum number $n = \infty$. Thus, the series limit is the shortest wavelength in the series.

If a jump is upward into the nonquantized portion of Fig. 39-18, the electron's energy is no longer given by Eq. 39-33 because the electron is no longer trapped in the atom. That is, the hydrogen atom has been *ionized*, meaning that the electron has been removed to a distance so great that the Coulomb force on it from the nucleus is negligible. The atom can be ionized if it absorbs any wavelength greater than the series limit. The free electron then has only kinetic energy $K (= \frac{1}{2}mv^2)$, assuming a nonrelativistic situation).

Quantum Numbers for the Hydrogen Atom

Although the energies of the hydrogen atom states can be described by the single quantum number n , the wave functions describing these states require three quantum numbers, corresponding to the three dimensions in which the electron can move. The three quantum numbers, along with their names and the values that they may have, are shown in Table 39-2.

Table 39-2

Quantum Numbers for the Hydrogen Atom

Symbol	Name	Allowed Values
n	Principal quantum number	1, 2, 3, ...
ℓ	Orbital quantum number	0, 1, 2, ..., $n - 1$
m_ℓ	Orbital magnetic quantum number	$-\ell, -(\ell - 1), \dots, +(\ell - 1), +\ell$

**CHECKPOINT 5**

- (a) A group of quantum states of the hydrogen atom has $n = 5$. How many values of ℓ are possible for states in this group? (b) A subgroup of hydrogen atom states in the $n = 5$ group has $\ell = 3$. How many values of m_ℓ are possible for states in this subgroup?

Each set of quantum numbers (n, ℓ, m_ℓ) identifies the wave function of a particular quantum state. The quantum number n , called the **principal quantum number**, appears in Eq. 39-33 for the energy of the state. The **orbital quantum number** ℓ is a measure of the magnitude of the angular momentum associated with the quantum state. The **orbital magnetic quantum number** m_ℓ is related to the orientation in space of this angular momentum vector. The restrictions on the values of the quantum numbers for the hydrogen atom, as listed in Table 39-2, are not arbitrary but come out of the solution to Schrödinger's equation. Note that for the ground state ($n = 1$), the restrictions require that $\ell = 0$ and $m_\ell = 0$. That is, the hydrogen atom in its ground state has zero angular momentum, which is not predicted by Eq. 39-24 in the Bohr model.

The Wave Function of the Hydrogen Atom's Ground State

The wave function for the ground state of the hydrogen atom, as obtained by solving the three-dimensional Schrödinger equation and normalizing the result, is

$$\psi(r) = \frac{1}{\sqrt{\pi a^{3/2}}} e^{-r/a} \quad (\text{ground state}), \quad (39-39)$$

where a is the Bohr radius (Eq. 39-29). This radius is loosely taken to be the effective radius of a hydrogen atom and turns out to be a convenient unit of length for other situations involving atomic dimensions.

As with other wave functions, $\psi(r)$ in Eq. 39-39 does not have physical meaning but $\psi^2(r)$ does, being the probability density—the probability per unit volume—that the electron can be detected. Specifically, $\psi^2(r) dV$ is the probability that the electron can be detected in any given (infinitesimal) volume element dV located at radius r from the center of the atom:

$$\begin{pmatrix} \text{probability of detection} \\ \text{in volume } dV \\ \text{at radius } r \end{pmatrix} = \begin{pmatrix} \text{volume probability} \\ \text{density } \psi^2(r) \\ \text{at radius } r \end{pmatrix} (\text{volume } dV). \quad (39-40)$$

Because $\psi^2(r)$ here depends only on r , it makes sense to choose, as a volume element dV , the volume between two concentric spherical shells whose radii are r and $r + dr$. That is, we take the volume element dV to be

$$dV = (4\pi r^2) dr, \quad (39-41)$$

in which $4\pi r^2$ is the surface area of the inner shell and dr is the radial distance between the two shells. Then, combining Eqs. 39-39, 39-40, and 39-41 gives us

$$\begin{pmatrix} \text{probability of detection} \\ \text{in volume } dV \\ \text{at radius } r \end{pmatrix} = \psi^2(r) dV = \frac{4}{a^3} e^{-2r/a} r^2 dr. \quad (39-42)$$

Describing the probability of detecting an electron is easier if we work with a **radial probability density** $P(r)$ instead of a volume probability density $\psi^2(r)$. This $P(r)$ is a linear probability density such that

$$\begin{pmatrix} \text{radial probability} \\ \text{density } P(r) \\ \text{at radius } r \end{pmatrix} \left(\frac{\text{radial}}{\text{width } dr} \right) = \begin{pmatrix} \text{volume probability} \\ \text{density } \psi^2(r) \\ \text{at radius } r \end{pmatrix} (\text{volume } dV)$$

$$\text{or} \quad P(r) dr = \psi^2(r) dV. \quad (39-43)$$

Substituting for $\psi^2(r) dV$ from Eq. 39-42, we obtain

$$P(r) = \frac{4}{a^3} r^2 e^{-2r/a} \quad (\text{radial probability density, hydrogen atom ground state}). \quad (39-44)$$

39-9 SCHRÖDINGER'S EQUATION AND THE HYDROGEN ATOM

1103

Figure 39-19 is a plot of Eq. 39-44. The area under the plot is unity; that is,

$$\int_0^{\infty} P(r) dr = 1. \quad (39-45)$$

This equation states that in a hydrogen atom, the electron must be *somewhere* in the space surrounding the nucleus.

The triangular marker on the horizontal axis of Fig. 39-19 is located one Bohr radius from the origin. The graph tells us that in the ground state of the hydrogen atom, the electron is most likely to be found at about this distance from the center of the atom.

Figure 39-19 conflicts sharply with the popular view that electrons in atoms follow well-defined orbits like planets moving around the Sun. *This popular view, however familiar, is incorrect.* Figure 39-19 shows us all that we can ever know about the location of the electron in the ground state of the hydrogen atom. The appropriate question is not “When will the electron arrive at such-and-such a point?” but “What are the odds that the electron will be detected in a small volume centered on such-and-such a point?” Figure 39-20, which we call a dot plot, suggests the probabilistic nature of the wave function and provides a useful mental model of the hydrogen atom in its ground state. Think of the atom in this state as a fuzzy ball with no sharply defined boundary and no hint of orbits.

It is not easy for a beginner to envision subatomic particles in this probabilistic way. The difficulty is our natural impulse to regard an electron as something like a tiny jelly bean, located at certain places at certain times and following a well-defined path. Electrons and other subatomic particles simply do not behave in this way.

The energy of the ground state, found by putting $n = 1$ in Eq. 39-33, is $E_1 = -13.60 \text{ eV}$. The wave function of Eq. 39-39 results if you solve Schrödinger's equation with this value of the energy. Actually, you can find a solution of Schrödinger's equation for *any* value of the energy—say, $E = -11.6 \text{ eV}$ or -14.3 eV . This may suggest that the energies of the hydrogen atom states are not quantized—but we know that they are.

The puzzle was solved when physicists realized that such solutions of Schrödinger's equation are not physically acceptable because they yield increasingly large values as $r \rightarrow \infty$. These “wave functions” tell us that the electron is more likely to be found very far from the nucleus rather than closer to it, which makes no sense. We get rid of these unwanted solutions by imposing what is called a **boundary condition**, in which we agree to accept only solutions of Schrödinger's equation for which $\psi(r) \rightarrow 0$ as $r \rightarrow \infty$; that is, we agree to deal only with *confined* electrons. With this restriction, the solutions of Schrödinger's equation form a discrete set, with quantized energies given by Eq. 39-33.

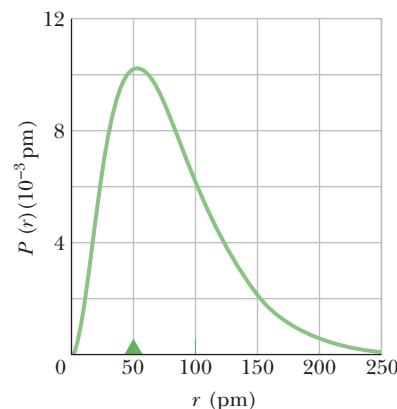


Fig. 39-19 A plot of the radial probability density $P(r)$ for the ground state of the hydrogen atom. The triangular marker is located at one Bohr radius from the origin, and the origin represents the center of the atom.

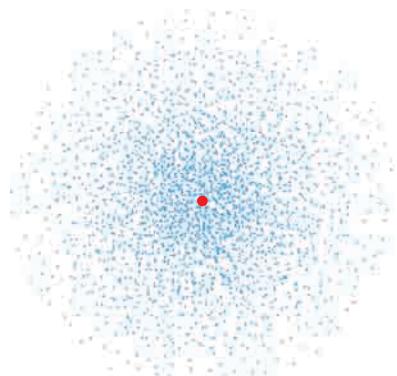


Fig. 39-20 A “dot plot” showing the volume probability density $\psi^2(r)$ —not the radial probability density $P(r)$ —for the ground state of the hydrogen atom. The density of dots drops exponentially with increasing distance from the nucleus, which is represented here by a red spot.

Sample Problem

Radial probability density for the electron in a hydrogen atom

Show that the radial probability density for the ground state of the hydrogen atom has a maximum at $r = a$.

KEY IDEAS

- (1) The radial probability density for a ground-state hydrogen atom is given by Eq. 39-44,

$$P(r) = \frac{4}{a^3} r^2 e^{-2r/a}.$$

- (2) To find the maximum (or minimum) of any function, we must differentiate the function and set the result equal to zero.

Calculation: If we differentiate $P(r)$ with respect to r , using derivative 7 of Appendix E and the chain rule for differentiating products, we get

$$\begin{aligned} \frac{dP}{dr} &= \frac{4}{a^3} r^2 \left(\frac{-2}{a} \right) e^{-2r/a} + \frac{4}{a^3} 2r e^{-2r/a} \\ &= \frac{8r}{a^3} e^{-2r/a} - \frac{8r^2}{a^4} e^{-2r/a} \\ &= \frac{8}{a^4} r(a - r) e^{-2r/a}. \end{aligned}$$

If we set the right side equal to zero, we obtain an equation that is true if $r = a$, so that the term $(a - r)$ in the middle of the equation is zero. In other words, dP/dr is equal to zero

when $r = a$. (Note that we also have $dP/dr = 0$ at $r = 0$ and at $r = \infty$. However, these conditions correspond to a *minimum* in $P(r)$, as you can see in Fig. 39-19.)

Sample Problem

Probability of detection of the electron in a hydrogen atom

It can be shown that the probability $p(r)$ that the electron in the ground state of the hydrogen atom will be detected inside a sphere of radius r is given by

$$p(r) = 1 - e^{-2x}(1 + 2x + 2x^2),$$

in which x , a dimensionless quantity, is equal to r/a . Find r for $p(r) = 0.90$.

KEY IDEA

There is no guarantee of detecting the electron at any particular radial distance r from the center of the hydrogen atom. However, with the given function, we can calculate the probability that the electron will be detected *somewhere* within a sphere of radius r .

Calculation: We seek the radius of a sphere for which $p(r) = 0.90$. Substituting that value in the expression for $p(r)$, we have

$$0.90 = 1 - e^{-2x}(1 + 2x + 2x^2)$$

$$\text{or} \quad 10e^{-2x}(1 + 2x + 2x^2) = 1.$$

We must find the value of x that satisfies this equality. It is not possible to solve explicitly for x , but an equation solver on a calculator yields $x = 2.66$. This means that the radius of a sphere within which the electron will be detected 90% of the time is $2.66a$. Mark this position on the horizontal axis of Fig. 39-19. The area under the curve from $r = 0$ to $r = 2.66a$ gives the probability of detection in that range and is 90% of the total area under the curve.

Sample Problem

Light emission from a hydrogen atom

- (a) What is the wavelength of light for the least energetic photon emitted in the Lyman series of the hydrogen atom spectrum lines?

KEY IDEAS

- (1) For any series, the transition that produces the least energetic photon is the transition between the home-base level that defines the series and the level immediately above it.
- (2) For the Lyman series, the home-base level is at $n = 1$ (Fig. 39-18b). Thus, the transition that produces the least energetic photon is the transition from the $n = 2$ level to the $n = 1$ level.

Calculations: From Eq. 39-33 the energy difference is

$$\Delta E = E_2 - E_1 = -(13.60 \text{ eV}) \left(\frac{1}{2^2} - \frac{1}{1^2} \right) = 10.20 \text{ eV}.$$

Then from Eq. 39-6 ($\Delta E = hf$), with c/λ replacing f , we have

$$\lambda = \frac{hc}{\Delta E} = \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s})}{(10.20 \text{ eV})(1.60 \times 10^{-19} \text{ J/eV})}$$

$$= 1.22 \times 10^{-7} \text{ m} = 122 \text{ nm.} \quad (\text{Answer})$$

Light with this wavelength is in the ultraviolet range.

- (b) What is the wavelength of the series limit for the Lyman series?

KEY IDEA

The series limit corresponds to a jump between the home-base level ($n = 1$ for the Lyman series) and the level at the limit $n = \infty$.

Calculations: Now that we have identified the values of n for the transition, we could proceed as in (a) to find the corresponding wavelength λ . Instead, let's use a more direct procedure. From Eq. 39-36, we find

$$\begin{aligned} \frac{1}{\lambda} &= R \left(\frac{1}{n_{\text{low}}^2} - \frac{1}{n_{\text{high}}^2} \right) \\ &= 1.097 373 \times 10^7 \text{ m}^{-1} \left(\frac{1}{1^2} - \frac{1}{\infty^2} \right), \end{aligned}$$

which yields

$$\lambda = 9.11 \times 10^{-8} \text{ m} = 91.1 \text{ nm.} \quad (\text{Answer})$$

Light with this wavelength is also in the ultraviolet range.



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39-9 SCHRÖDINGER'S EQUATION AND THE HYDROGEN ATOM

1105

Table 39-3

Quantum Numbers for Hydrogen Atom States with $n = 2$

n	ℓ	m_ℓ
2	0	0
2	1	+1
2	1	0
2	1	-1

Hydrogen Atom States with $n = 2$

According to the requirements of Table 39-2, there are four states of the hydrogen atom with $n = 2$; their quantum numbers are listed in Table 39-3. Consider first the state with $n = 2$ and $\ell = m_\ell = 0$; its probability density is represented by the dot plot of Fig. 39-21. Note that this plot, like the plot for the ground state shown in Fig. 39-20, is spherically symmetric. That is, in a spherical coordinate system like that defined in Fig. 39-22, the probability density is a function of the radial coordinate r only and is independent of the angular coordinates θ and ϕ .

It turns out that all quantum states with $\ell = 0$ have spherically symmetric wave functions. This is reasonable because the quantum number ℓ is a measure of the angular momentum associated with a given state. If $\ell = 0$, the angular momentum is also zero, which requires that the probability density representing the state have no preferred axis of symmetry.

Dot plots of ψ^2 for the three states with $n = 2$ and $\ell = 1$ are shown in Fig. 39-23. The probability densities for the states with $m_\ell = +1$ and $m_\ell = -1$ are identical. Although these plots are symmetric about the z axis, they are *not* spherically symmetric. That is, the probability densities for these three states are functions of both r and the angular coordinate θ .

Here is a puzzle: What is there about the hydrogen atom that establishes the axis of symmetry that is so obvious in Fig. 39-23? The answer: *absolutely nothing*.

The solution to this puzzle comes about when we realize that all three states shown in Fig. 39-23 have the same energy. Recall that the energy of a state, given by Eq. 39-33, depends only on the principal quantum number n and is independent of ℓ and m_ℓ . In fact, for an *isolated* hydrogen atom there is no way to differentiate experimentally among the three states of Fig. 39-23.

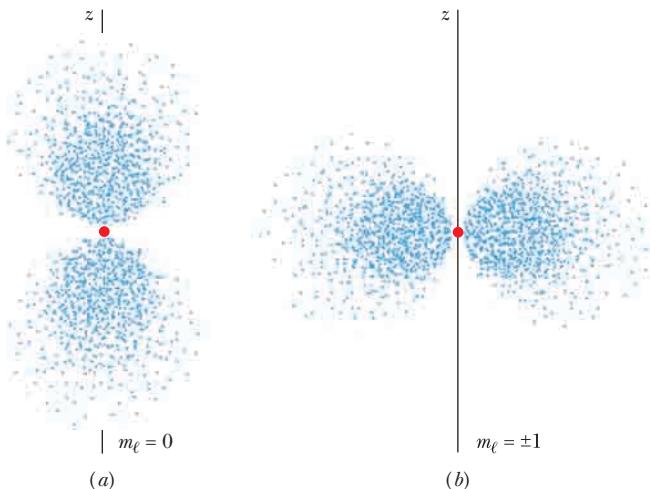


Fig. 39-23 Dot plots of the volume probability density $\psi^2(r, \theta)$ for the hydrogen atom in states with $n = 2$ and $\ell = 1$. (a) Plot for $m_\ell = 0$. (b) Plot for $m_\ell = +1$ and $m_\ell = -1$. Both plots show that the probability density is symmetric about the z axis.

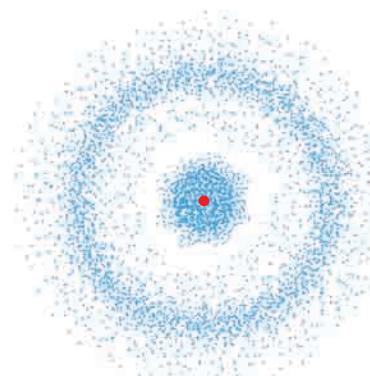


Fig. 39-21 A dot plot showing the volume probability density $\psi^2(r)$ for the hydrogen atom in the quantum state with $n = 2$, $\ell = 0$, and $m_\ell = 0$. The plot has spherical symmetry about the central nucleus. The gap in the dot density pattern marks a spherical surface over which $\psi^2(r) = 0$.

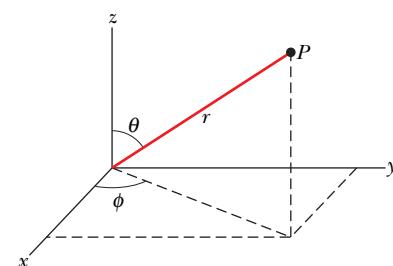


Fig. 39-22 The relationship between the coordinates x , y , and z of the rectangular coordinate system and the coordinates r , θ , and ϕ of the spherical coordinate system. The latter are more appropriate for analyzing situations involving spherical symmetry, such as the hydrogen atom.

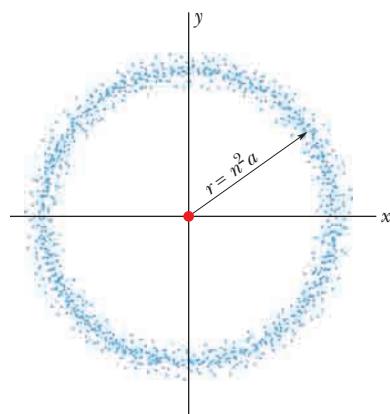


Fig. 39-24 A dot plot of the radial probability density $P(r)$ for the hydrogen atom in a quantum state with a relatively large principal quantum number—namely, $n = 45$ —and angular momentum quantum number $\ell = n - 1 = 44$. The dots lie close to the xy plane, the ring of dots suggesting a classical electron orbit.

If we add the volume probability densities for the three states for which $n = 2$ and $\ell = 1$, the combined probability density turns out to be spherically symmetrical, with no unique axis. One can, then, think of the electron as spending one-third of its time in each of the three states of Fig. 39-23, and one can think of the weighted sum of the three independent wave functions as defining a spherically symmetric **subshell** specified by the quantum numbers $n = 2$, $\ell = 1$. The individual states will display their separate existence only if we place the hydrogen atom in an external electric or magnetic field. The three states of the $n = 2$, $\ell = 1$ subshell will then have different energies, and the field direction will establish the necessary symmetry axis.

The $n = 2$, $\ell = 0$ state, whose volume probability density is shown in Fig. 39-21, also has the same energy as each of the three states of Fig. 39-23. We can view all four states whose quantum numbers are listed in Table 39-3 as forming a spherically symmetric **shell** specified by the single quantum number n . The importance of shells and subshells will become evident in Chapter 40, where we discuss atoms having more than one electron.

To round out our picture of the hydrogen atom, we display in Fig. 39-24 a dot plot of the *radial* probability density for a hydrogen atom state with a relatively high quantum number ($n = 45$) and the highest orbital quantum number that the restrictions of Table 39-2 permit ($\ell = n - 1 = 44$). The probability density forms a ring that is symmetrical about the z axis and lies very close to the xy plane. The mean radius of the ring is $n^2 a$, where a is the Bohr radius. This mean radius is more than 2000 times the effective radius of the hydrogen atom in its ground state.

Figure 39-24 suggests the electron orbit of classical physics—it resembles the circular orbit of a planet around a star. Thus, we have another illustration of Bohr's correspondence principle—namely, that at large quantum numbers the predictions of quantum mechanics merge smoothly with those of classical physics. Imagine what a dot plot like that of Figure 39-24 would look like for *really* large values of n and ℓ —say, $n = 1000$ and $\ell = 999$.

REVIEW & SUMMARY

The Confinement Principle The **confinement principle** applies to waves of all kinds, including waves on a string and the matter waves of quantum physics. It states that confinement leads to quantization—that is, to the existence of discrete states with certain energies. States with intermediate values of energy are disallowed.

An Electron in an Infinite Potential Well An infinite potential well is a device for confining an electron. From the confinement principle we expect that the matter wave representing a trapped electron can exist only in a set of discrete states. For a one-dimensional infinite potential well, the energies associated with these *quantum states* are

$$E_n = \left(\frac{h^2}{8mL^2} \right) n^2, \quad \text{for } n = 1, 2, 3, \dots, \quad (39-4)$$

in which L is the width of the well and n is a **quantum number**. If the electron is to change from one state to another, its energy must change by the amount

$$\Delta E = E_{\text{high}} - E_{\text{low}}, \quad (39-5)$$

where E_{high} is the higher energy and E_{low} is the lower energy. If the change is done by photon absorption or emission, the energy of the photon must be

$$hf = \Delta E = E_{\text{high}} - E_{\text{low}}. \quad (39-6)$$

The **wave functions** associated with the quantum states are

$$\psi_n(x) = A \sin\left(\frac{n\pi}{L} x\right), \quad \text{for } n = 1, 2, 3, \dots \quad (39-10)$$

The probability density $\psi_n^2(x)$ for an allowed state has the physical meaning that $\psi_n^2(x) dx$ is the probability that the electron will be detected in the interval between x and $x + dx$. For an electron in an infinite well, the probability densities are

$$\psi_n^2(x) = A^2 \sin^2\left(\frac{n\pi}{L} x\right), \quad \text{for } n = 1, 2, 3, \dots \quad (39-12)$$

At high quantum numbers n , the electron tends toward classical behavior in that it tends to occupy all parts of the well with equal probability. This transition from quantum to classical physics is known as the **correspondence principle**.

Normalization and Zero-Point Energy The amplitude A^2 in Eq. 39-12 can be found from the **normalizing equation**,

$$\int_{-\infty}^{+\infty} \psi_n^2(x) dx = 1, \quad (39-14)$$

which asserts that the electron must be *somewhere* within the well because the probability 1 implies certainty.

From Eq. 39-4 we see that the lowest permitted energy for the electron is not zero but the energy that corresponds to $n = 1$. This lowest energy is called the **zero-point energy** of the electron–well system.

An Electron in a Finite Potential Well A finite potential well is one for which the potential energy of an electron inside the well is less than that for one outside the well by a finite amount U_0 . The wave function for an electron trapped in such a well extends into the walls of the well.

Two- and Three-Dimensional Electron Traps The quantized energies for an electron trapped in a two-dimensional infinite potential well that forms a rectangular corral are

$$E_{nx,ny} = \frac{\hbar^2}{8m} \left(\frac{n_x^2}{L_x^2} + \frac{n_y^2}{L_y^2} \right), \quad (39-20)$$

where n_x is a quantum number for which the electron's matter wave fits in well width L_x and n_y is a quantum number for which the electron's matter wave fits in well width L_y . Similarly, the energies for an electron trapped in a three-dimensional infinite potential well that forms a rectangular box are

$$E_{nx,ny,nz} = \frac{\hbar^2}{8m} \left(\frac{n_x^2}{L_x^2} + \frac{n_y^2}{L_y^2} + \frac{n_z^2}{L_z^2} \right). \quad (39-21)$$

Here n_z is a third quantum number, one for which the matter wave fits in well width L_z .

The Hydrogen Atom Both the (incorrect) Bohr model of the hydrogen atom and the (correct) application of Schrödinger's equation to this atom give the quantized energy levels of the atom as

$$E_n = -\frac{me^4}{8\varepsilon_0^2 h^2} \frac{1}{n^2} = -\frac{13.60 \text{ eV}}{n^2}, \quad (39-32, 39-33)$$

for $n = 1, 2, 3, \dots$

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QUESTIONS

1 Three electrons are trapped in three different one-dimensional infinite potential wells of widths (a) 50 pm, (b) 200 pm, and (c) 100 pm. Rank the electrons according to their ground-state energies, greatest first.

2 Is the ground-state energy of a proton trapped in a one-dimensional infinite potential well greater than, less than, or equal to that of an electron trapped in the same potential well?

3 An electron is trapped in a one-dimensional infinite potential well in a state with $n = 17$. How many points of (a) zero probability and (b) maximum probability does its matter wave have?

4 Figure 39-25 shows three infinite potential wells, each on an x axis. Without written calculation, determine the wave function ψ for a ground-state electron trapped in each well.

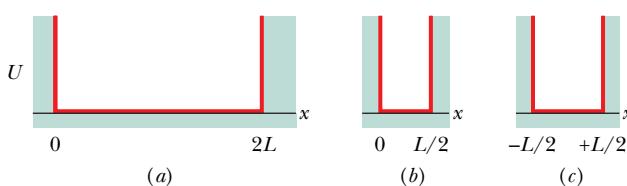


Fig. 39-25 Question 4.

From this we find that if the atom makes a transition between any two energy levels as a result of having emitted or absorbed light, the wavelength of the light is given by

$$\frac{1}{\lambda} = R \left(\frac{1}{n_{\text{low}}^2} - \frac{1}{n_{\text{high}}^2} \right), \quad (39-36)$$

where

$$R = \frac{me^4}{8\varepsilon_0^2 h^3 c} = 1.097 373 \times 10^7 \text{ m}^{-1} \quad (39-37)$$

is the **Rydberg constant**.

The **radial probability density** $P(r)$ for a state of the hydrogen atom is defined so that $P(r) dr$ is the probability that the electron will be detected somewhere in the space between two concentric shells of radii r and $r + dr$ centered on the atom's nucleus. For the hydrogen atom's ground state,

$$P(r) = \frac{4}{a^3} r^2 e^{-2r/a}, \quad (39-44)$$

in which a , the **Bohr radius**, is a length unit equal to 52.92 pm. Figure 39-19 is a plot of $P(r)$ for the ground state.

Figures 39-21 and 39-23 represent the volume probability densities (not the *radial* probability densities) for the four hydrogen atom states with $n = 2$. The plot of Fig. 39-21 ($n = 2, \ell = 0, m_\ell = 0$) is spherically symmetric. The plots of Fig. 39-23 ($n = 2, \ell = 1, m_\ell = 0, +1, -1$) are symmetric about the z axis but, when added together, are also spherically symmetric.

All four states with $n = 2$ have the same energy and may be usefully regarded as constituting a **shell**, identified as the $n = 2$ shell. The three states of Fig. 39-23, taken together, may be regarded as constituting the $n = 2, \ell = 1$ **subshell**. It is not possible to separate the four $n = 2$ states experimentally unless the hydrogen atom is placed in an electric or magnetic field, which permits the establishment of a definite symmetry axis.

5 A proton and an electron are trapped in identical one-dimensional infinite potential wells; each particle is in its ground state. At the center of the wells, is the probability density for the proton greater than, less than, or equal to that of the electron?

6 If you double the width of a one-dimensional infinite potential well, (a) is the energy of the ground state of the trapped electron multiplied by 4, 2, $\frac{1}{2}$, $\frac{1}{4}$, or some other number? (b) Are the energies of the higher energy states multiplied by this factor or by some other factor, depending on their quantum number?

7 If you wanted to use the idealized trap of Fig. 39-1 to trap a positron, would you need to change (a) the geometry of the trap, (b) the electric potential of the central cylinder, or (c) the electric potentials of the two semi-infinite end cylinders? (A positron has the same mass as an electron but is positively charged.)

8 An electron is trapped in a finite potential well that is deep enough to allow the electron to exist in a state with $n = 4$. How many points of (a) zero probability and (b) maximum probability does its matter wave have within the well?

9 An electron that is trapped in a one-dimensional infinite potential well of width L is excited from the ground state to the first ex-

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1108 CHAPTER 39 MORE ABOUT MATTER WAVES

cited state. Does the excitation increase, decrease, or have no effect on the probability of detecting the electron in a small length of the x axis (a) at the center of the well and (b) near one of the well walls?

10 An electron, trapped in a finite potential energy well such as that of Fig. 39-7, is in its state of lowest energy. Are (a) its de Broglie wavelength, (b) the magnitude of its momentum, and (c) its energy greater than, the same as, or less than they would be if the potential well were infinite, as in Fig. 39-2?

11 From a visual inspection of Fig. 39-8, rank the quantum numbers of the three quantum states according to the de Broglie wavelength of the electron, greatest first.

12 You want to modify the finite potential well of Fig. 39-7 to allow its trapped electron to exist in more than four quantum states. Could you do so by making the well (a) wider or narrower, (b) deeper or shallower?

13 A hydrogen atom is in the third excited state. To what state (give the quantum number n) should it jump to (a) emit light with the longest possible wavelength, (b) emit light with the shortest possible wavelength, and (c) absorb light with the longest possible wavelength?

14 Figure 39-26 indicates the lowest energy levels (in electron-volts) for five situations in which an electron is trapped in a one-dimensional infinite potential well. In wells B , C , D , and E , the electron is in the ground state. We shall excite the electron in well A to the fourth excited state (at 25 eV). The electron can then de-excite to the ground state by emitting one or more photons, corresponding to one long jump or several short jumps. Which photon *emis-*

sion energies of this de-excitation match a photon *absorption* energy (from the ground state) of the other four electrons? Give the n values.

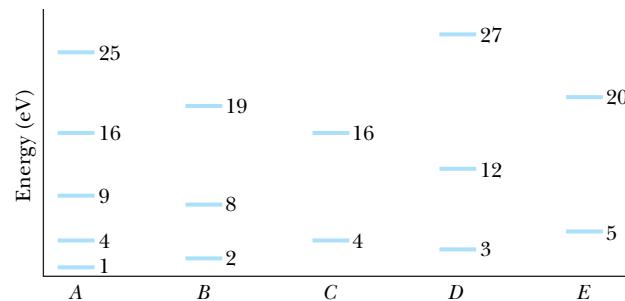


Fig. 39-26 Question 14.

15 Table 39-4 lists the quantum numbers for five proposed hydrogen atom states. Which of them are not possible?

Table 39-4

	n	ℓ	m_ℓ
(a)	3	2	0
(b)	2	3	1
(c)	4	3	-4
(d)	5	5	0
(e)	5	3	-2



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>

sec. 39-3 Energies of a Trapped Electron

•1 An electron in a one-dimensional infinite potential well of length L has ground-state energy E_1 . The length is changed to L' so that the new ground-state energy is $E'_1 = 0.500E_1$. What is the ratio L'/L ?

•2 What is the ground-state energy of (a) an electron and (b) a proton if each is trapped in a one-dimensional infinite potential well that is 200 pm wide?

•3 The ground-state energy of an electron trapped in a one-dimensional infinite potential well is 2.6 eV. What will this quantity be if the width of the potential well is doubled?

•4 An electron, trapped in a one-dimensional infinite potential well 250 pm wide, is in its ground state. How much energy must it absorb if it is to jump up to the state with $n = 4$?

•5 What must be the width of a one-dimensional infinite potential well if an electron trapped in it in the $n = 3$ state is to have an energy of 4.7 eV?

•6 A proton is confined to a one-dimensional infinite potential well 100 pm wide. What is its ground-state energy?

•7 Consider an atomic nucleus to be equivalent to a one-dimensional infinite potential well with $L = 1.4 \times 10^{-14}$ m, a typical nuclear diameter. What would be the ground-state energy of an electron if it were trapped in such a potential well? (Note: Nuclei do not contain electrons.)

•8 An electron is trapped in a one-dimensional infinite well and is in its first excited state. Figure 39-27 indicates the five longest wavelengths of light that the electron could absorb in transitions from this initial state via a single photon absorption: $\lambda_a = 80.78$ nm, $\lambda_b = 33.66$ nm, $\lambda_c = 19.23$ nm, $\lambda_d = 12.62$ nm, and $\lambda_e = 8.98$ nm. What is the width of the potential well?



Fig. 39-27 Problem 8.

•9 Suppose that an electron trapped in a one-dimensional infinite well of width 250 pm is excited from its first excited state to its third excited state. (a) What energy must be transferred to the electron for this quantum jump? The electron then de-excites back to its ground

state by emitting light. In the various possible ways it can do this, what are the (b) shortest, (c) second shortest, (d) longest, and (e) second longest wavelengths that can be emitted? (f) Show the various possible ways on an energy-level diagram. If light of wavelength 29.4 nm happens to be emitted, what are the (g) longest and (h) shortest wavelength that can be emitted afterwards?

••10 An electron is trapped in a one-dimensional infinite potential well. For what (a) higher quantum number and (b) lower quantum number is the corresponding energy difference equal to the energy difference ΔE_{43} between the levels $n = 4$ and $n = 3$? (c) Show that no pair of adjacent levels has an energy difference equal to $2\Delta E_{43}$.

••11 An electron is trapped in a one-dimensional infinite potential well. For what (a) higher quantum number and (b) lower quantum number is the corresponding energy difference equal to the energy of the $n = 5$ level? (c) Show that no pair of adjacent levels has an energy difference equal to the energy of the $n = 6$ level.

••12 An electron is trapped in a one-dimensional infinite well of width 250 pm and is in its ground state. What are the (a) longest, (b) second longest, and (c) third longest wavelengths of light that can excite the electron from the ground state via a single photon absorption?

sec. 39-4 Wave Functions of a Trapped Electron

••13 A one-dimensional infinite well of length 200 pm contains an electron in its third excited state. We position an electron-detector probe of width 2.00 pm so that it is centered on a point of maximum probability density. (a) What is the probability of detection by the probe? (b) If we insert the probe as described 1000 times, how many times should we expect the electron to materialize on the end of the probe (and thus be detected)?

••14 An electron is in a certain energy state in a one-dimensional, infinite potential well from $x = 0$ to $x = L = 200$ pm. The electron's probability density is zero at $x = 0.300L$, and $x = 0.400L$; it is not zero at intermediate values of x . The electron then jumps to the next lower energy level by emitting light. What is the change in the electron's energy?

••15 An electron is trapped in a one-dimensional infinite potential well that is 100 pm wide; the electron is in its ground state. What is the probability that you can detect the electron in an interval of width $\Delta x = 5.0$ pm centered at $x =$ (a) 25 pm, (b) 50 pm, and (c) 90 pm? (*Hint:* The interval Δx is so narrow that you can take the probability density to be constant within it.)

••16 A particle is confined to the one-dimensional infinite potential well of Fig. 39-2. If the particle is in its ground state, what is its probability of detection between (a) $x = 0$ and $x = 0.25L$, (b) $x = 0.75L$ and $x = L$, and (c) $x = 0.25L$ and $x = 0.75L$?

sec. 39-5 An Electron in a Finite Well

••17 An electron in the $n = 2$ state in the finite potential well of Fig. 39-7 absorbs 400 eV of energy from an external source. Using the energy-level diagram of Fig. 39-9, determine the electron's kinetic energy after this absorption, assuming that the electron moves to a position for which $x > L$.

••18 Figure 39-9 gives the energy levels for an electron trapped in a finite potential energy well 450 eV deep. If the electron is in the $n = 3$ state, what is its kinetic energy?

••19 Figure 39-28a shows the energy-level diagram for a finite, one-dimensional energy well that contains an electron. The non-

quantized region begins at $E_4 = 450.0$ eV. Figure 39-28b gives the absorption spectrum of the electron when it is in the ground state—it can absorb at the indicated wavelengths: $\lambda_a = 14.588$ nm and $\lambda_b = 4.8437$ nm and for any wavelength less than $\lambda_c = 2.9108$ nm. What is the energy of the first excited state?

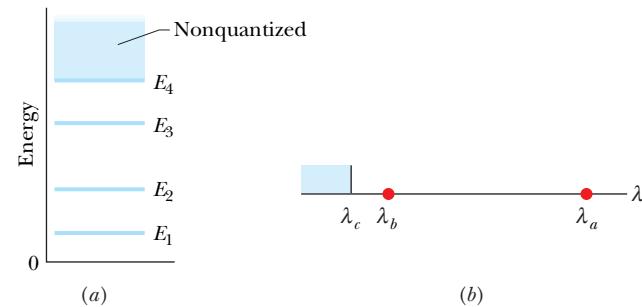


Fig. 39-28 Problem 19.

••20 Figure 39-29a shows a thin tube in which a finite potential trap has been set up where $V_2 = 0$ V. An electron is shown traveling rightward toward the trap, in a region with a voltage of $V_1 = -9.00$ V, where it has a kinetic energy of 2.00 eV. When the electron enters the trap region, it can become trapped if it gets rid of enough energy by emitting a photon. The energy levels of the electron within the trap are $E_1 = 1.0$, $E_2 = 2.0$, and $E_3 = 4.0$ eV, and the nonquantized region begins at $E_4 = 9.0$ eV as shown in the energy-level diagram of Fig. 39-29b. What is the smallest energy (eV) such a photon can have?

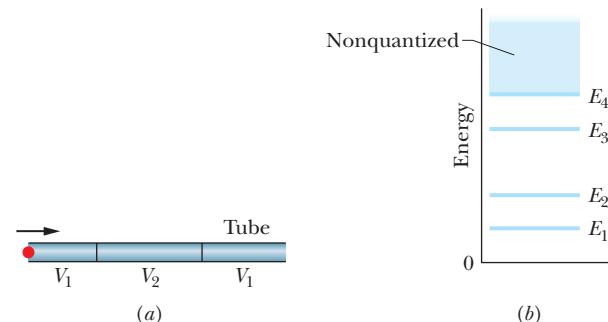


Fig. 39-29 Problem 20.

••21 (a) Show that for the region $x > L$ in the finite potential well of Fig. 39-7, $\psi(x) = De^{kx}$ is a solution of Schrödinger's equation in its one-dimensional form, where D is a constant and k is positive. (b) On what basis do we find this mathematically acceptable solution to be physically unacceptable?

sec. 39-7 Two- and Three-Dimensional Electron Traps

••22 An electron is contained in the rectangular corral of Fig. 39-13, with widths $L_x = 800$ pm and $L_y = 1600$ pm. What is the electron's ground-state energy?

••23 An electron is contained in the rectangular box of Fig. 39-14, with widths $L_x = 800$ pm, $L_y = 1600$ pm, and $L_z = 390$ pm. What is the electron's ground-state energy?

••24 Figure 39-30 shows a two-dimensional, infinite-potential well lying in an xy plane that contains an electron. We probe

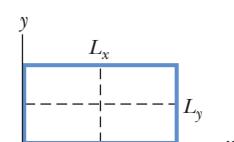


Fig. 39-30
Problem 24.

1110 CHAPTER 39 MORE ABOUT MATTER WAVES

for the electron along a line that bisects L_x and find three points at which the detection probability is maximum. Those points are separated by 2.00 nm. Then we probe along a line that bisects L_y and find five points at which the detection probability is maximum. Those points are separated by 3.00 nm. What is the energy of the electron?

- 25 GO** The two-dimensional, infinite corral of Fig. 39-31 is square, with edge length $L = 150$ pm. A square probe is centered at xy coordinates $(0.200L, 0.800L)$ and has an x width of 5.00 pm and a y width of 5.00 pm. What is the probability of detection if the electron is in the $E_{1,3}$ energy state?

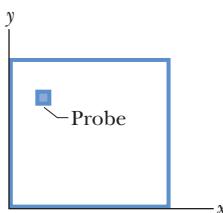


Fig. 39-31

Problem 25.

- 26** A rectangular corral of widths $L_x = L$ and $L_y = 2L$ contains an electron. What multiple of $h^2/8mL^2$, where m is the electron mass, gives (a) the energy of the electron's ground state, (b) the energy of its first excited state, (c) the energy of its lowest degenerate states, and (d) the difference between the energies of its second and third excited states?

- 27 SSM WWW** An electron (mass m) is contained in a rectangular corral of widths $L_x = L$ and $L_y = 2L$. (a) How many different frequencies of light could the electron emit or absorb if it makes a transition between a pair of the lowest five energy levels? What multiple of $h/8mL^2$ gives the (b) lowest, (c) second lowest, (d) third lowest, (e) highest, (f) second highest, and (g) third highest frequency?

- 28 GO** A cubical box of widths $L_x = L_y = L_z = L$ contains an electron. What multiple of $h^2/8mL^2$, where m is the electron mass, is (a) the energy of the electron's ground state, (b) the energy of its second excited state, and (c) the difference between the energies of its second and third excited states? How many degenerate states have the energy of (d) the first excited state and (e) the fifth excited state?

- 29** An electron (mass m) is contained in a cubical box of widths $L_x = L_y = L_z$. (a) How many different frequencies of light could the electron emit or absorb if it makes a transition between a pair of the lowest five energy levels? What multiple of $h/8mL^2$ gives the (b) lowest, (c) second lowest, (d) third lowest, (e) highest, (f) second highest, and (g) third highest frequency?

- 30** An electron is in the ground state in a two-dimensional, square, infinite potential well with edge lengths L . We will probe for it in a square of area 400 pm 2 that is centered at $x = L/8$ and $y = L/8$. The probability of detection turns out to be 4.5×10^{-8} . What is edge length L ?

sec. 39-9 Schrödinger's Equation and the Hydrogen Atom

- 31 SSM** What is the ratio of the shortest wavelength of the Balmer series to the shortest wavelength of the Lyman series?

- 32** An atom (not a hydrogen atom) absorbs a photon whose associated wavelength is 375 nm and then immediately emits a photon whose associated wavelength is 580 nm. How much net energy is absorbed by the atom in this process?

- 33** What are the (a) energy, (b) magnitude of the momentum, and (c) wavelength of the photon emitted when a hydrogen atom undergoes a transition from a state with $n = 3$ to a state with $n = 1$?

- 34** Calculate the radial probability density $P(r)$ for the hydrogen atom in its ground state at (a) $r = 0$, (b) $r = a$, and (c) $r = 2a$, where a is the Bohr radius.

- 35** For the hydrogen atom in its ground state, calculate (a) the probability density $\psi^2(r)$ and (b) the radial probability density $P(r)$ for $r = a$, where a is the Bohr radius.

- 36** (a) What is the energy E of the hydrogen-atom electron whose probability density is represented by the dot plot of Fig. 39-21? (b) What minimum energy is needed to remove this electron from the atom?

- 37 SSM** A neutron with a kinetic energy of 6.0 eV collides with a stationary hydrogen atom in its ground state. Explain why the collision must be elastic—that is, why kinetic energy must be conserved. (*Hint:* Show that the hydrogen atom cannot be excited as a result of the collision.)

- 38** An atom (not a hydrogen atom) absorbs a photon whose associated frequency is 6.2×10^{14} Hz. By what amount does the energy of the atom increase?

- 39 SSM** Verify that Eq. 39-44, the radial probability density for the ground state of the hydrogen atom, is normalized. That is, verify that the following is true:

$$\int_0^\infty P(r) dr = 1$$

- 40** What are the (a) wavelength range and (b) frequency range of the Lyman series? What are the (c) wavelength range and (d) frequency range of the Balmer series?

- 41** What is the probability that an electron in the ground state of the hydrogen atom will be found between two spherical shells whose radii are r and $r + \Delta r$, (a) if $r = 0.500a$ and $\Delta r = 0.010a$ and (b) if $r = 1.00a$ and $\Delta r = 0.01a$, where a is the Bohr radius? (*Hint:* Δr is small enough to permit the radial probability density to be taken to be constant between r and $r + \Delta r$.)

- 42** A hydrogen atom, initially at rest in the $n = 4$ quantum state, undergoes a transition to the ground state, emitting a photon in the process. What is the speed of the recoiling hydrogen atom? (*Hint:* This is similar to the explosions of Chapter 9.)

- 43** In the ground state of the hydrogen atom, the electron has a total energy of -13.6 eV. What are (a) its kinetic energy and (b) its potential energy if the electron is one Bohr radius from the central nucleus?

- 44** A hydrogen atom in a state having a *binding energy* (the energy required to remove an electron) of 0.85 eV makes a transition to a state with an *excitation energy* (the difference between the energy of the state and that of the ground state) of 10.2 eV. (a) What is the energy of the photon emitted as a result of the transition? What are the (b) higher quantum number and (c) lower quantum number of the transition producing this emission?

- 45 SSM** The wave functions for the three states with the dot plots shown in Fig. 39-23, which have $n = 2$, $\ell = 1$, and $m_\ell = 0, +1$, and -1 , are

$$\begin{aligned}\psi_{210}(r, \theta) &= (1/4\sqrt{2\pi})(a^{-3/2})(r/a)e^{-r/2a} \cos \theta, \\ \psi_{21+1}(r, \theta) &= (1/8\sqrt{\pi})(a^{-3/2})(r/a)e^{-r/2a}(\sin \theta)e^{+i\phi}, \\ \psi_{21-1}(r, \theta) &= (1/8\sqrt{\pi})(a^{-3/2})(r/a)e^{-r/2a}(\sin \theta)e^{-i\phi},\end{aligned}$$

in which the subscripts on $\psi(r, \theta)$ give the values of the quantum numbers n, ℓ, m_ℓ and the angles θ and ϕ are defined in Fig. 39-22. Note that the first wave function is real but the others, which

involve the imaginary number i , are complex. Find the radial probability density $P(r)$ for (a) ψ_{210} and (b) ψ_{21+1} (same as for ψ_{21-1}). (c) Show that each $P(r)$ is consistent with the corresponding dot plot in Fig. 39-23. (d) Add the radial probability densities for ψ_{210} , ψ_{21+1} , and ψ_{21-1} and then show that the sum is spherically symmetric, depending only on r .

••46 Calculate the probability that the electron in the hydrogen atom, in its ground state, will be found between spherical shells whose radii are a and $2a$, where a is the Bohr radius.

••47 For what value of the principal quantum number n would the effective radius, as shown in a probability density dot plot for the hydrogen atom, be 1.0 mm? Assume that ℓ has its maximum value of $n - 1$. (*Hint:* See Fig. 39-24.)

••48 Light of wavelength 121.6 nm is emitted by a hydrogen atom. What are the (a) higher quantum number and (b) lower quantum number of the transition producing this emission? (c) What is the name of the series that includes the transition?

••49 How much work must be done to pull apart the electron and the proton that make up the hydrogen atom if the atom is initially in (a) its ground state and (b) the state with $n = 2$?

••50 Light of wavelength 102.6 nm is emitted by a hydrogen atom. What are the (a) higher quantum number and (b) lower quantum number of the transition producing this emission? (c) What is the name of the series that includes the transition?

••51 What is the probability that in the ground state of the hydrogen atom, the electron will be found at a radius greater than the Bohr radius?

••52 A hydrogen atom is excited from its ground state to the state with $n = 4$. (a) How much energy must be absorbed by the atom? Consider the photon energies that can be emitted by the atom as it de-excites to the ground state in the several possible ways. (b) How many different energies are possible; what are the (c) highest, (d) second highest, (e) third highest, (f) lowest, (g) second lowest, and (h) third lowest energies?

••53 SSM WWW Schrödinger's equation for states of the hydrogen atom for which the orbital quantum number ℓ is zero is

$$\frac{1}{r^2} \frac{d}{dr} \left(r^2 \frac{d\psi}{dr} \right) + \frac{8\pi^2 m}{h^2} [E - U(r)]\psi = 0.$$

Verify that Eq. 39-39, which describes the ground state of the hydrogen atom, is a solution of this equation.

••54 The wave function for the hydrogen-atom quantum state represented by the dot plot shown in Fig. 39-21, which has $n = 2$ and $\ell = m_\ell = 0$, is

$$\psi_{200}(r) = \frac{1}{4\sqrt{2\pi}} a^{-3/2} \left(2 - \frac{r}{a} \right) e^{-r/2a},$$

in which a is the Bohr radius and the subscript on $\psi(r)$ gives the values of the quantum numbers n , ℓ , m_ℓ . (a) Plot $\psi_{200}^2(r)$ and show that your plot is consistent with the dot plot of Fig. 39-21. (b) Show analytically that $\psi_{200}^2(r)$ has a maximum at $r = 4a$. (c) Find the radial probability density $P_{200}(r)$ for this state. (d) Show that

$$\int_0^\infty P_{200}(r) dr = 1$$

and thus that the expression above for the wave function $\psi_{200}(r)$ has been properly normalized.

••55 The radial probability density for the ground state of the hydrogen atom is a maximum when $r = a$, where a is the Bohr radius. Show that the average value of r , defined as

$$r_{\text{avg}} = \int P(r) r dr,$$

has the value $1.5a$. In this expression for r_{avg} , each value of $P(r)$ is weighted with the value of r at which it occurs. Note that the average value of r is greater than the value of r for which $P(r)$ is a maximum.

Additional Problems

56 Let ΔE_{adj} be the energy difference between two adjacent energy levels for an electron trapped in a one-dimensional infinite potential well. Let E be the energy of either of the two levels. (a) Show that the ratio $\Delta E_{\text{adj}}/E$ approaches the value $2/n$ at large values of the quantum number n . As $n \rightarrow \infty$, does (b) ΔE_{adj} , (c) E , or (d) $\Delta E_{\text{adj}}/E$ approach zero? (e) What do these results mean in terms of the correspondence principle?

57 An electron is trapped in a one-dimensional infinite potential well. Show that the energy difference ΔE between its quantum levels n and $n + 2$ is $(h^2/2mL^2)(n + 1)$.

58 As Fig. 39-8 suggests, the probability density for an electron in the region $0 < x < L$ for the finite potential well of Fig. 39-7 is sinusoidal, being given by $\psi^2(x) = B \sin^2 kx$, in which B is a constant. (a) Show that the wave function $\psi(x)$ that may be found from this equation is a solution of Schrödinger's equation in its one-dimensional form. (b) Find an expression for k that makes this true.

59 SSM As Fig. 39-8 suggests, the probability density for the region $x > L$ in the finite potential well of Fig. 39-7 drops off exponentially according to $\psi^2(x) = Ce^{-2kx}$, where C is a constant. (a) Show that the wave function $\psi(x)$ that may be found from this equation is a solution of Schrödinger's equation in its one-dimensional form. (b) Find an expression for k for this to be true.

60 An electron is confined to a narrow evacuated tube of length 3.0 m; the tube functions as a one-dimensional infinite potential well. (a) What is the energy difference between the electron's ground state and its first excited state? (b) At what quantum number n would the energy difference between adjacent energy levels be 1.0 eV—which is measurable, unlike the result of (a)? At that quantum number, (c) what multiple of the electron's rest energy would give the electron's total energy and (d) would the electron be relativistic?

61 (a) Show that the terms in Schrödinger's equation (Eq. 39-18) have the same dimensions. (b) What is the common SI unit for each of these terms?

62 (a) What is the wavelength of light for the least energetic photon emitted in the Balmer series of the hydrogen atom spectrum lines? (b) What is the wavelength of the series limit?

63 (a) For a given value of the principal quantum number n for a hydrogen atom, how many values of the orbital quantum number ℓ are possible? (b) For a given value of ℓ , how many values of the orbital magnetic quantum number m_ℓ are possible? (c) For a given value of n , how many values of m_ℓ are possible?

64 Verify that the combined value of the constants appearing in Eq. 39-32 is 13.6 eV.

40

ALL ABOUT ATOMS

40-1 WHAT IS PHYSICS?

In this chapter we continue with a primary goal of physics—discovering and understanding the properties of atoms. About 100 years ago, researchers struggled to find experiments that would prove the existence of atoms. Now we take their existence for granted and even have photographs (scanning tunneling microscope images) of atoms. We can drag them around on surfaces, such as to make the quantum corral shown in the photograph of Fig. 39-12. We can even hold an individual atom indefinitely in a trap (Fig. 40-1) so as to study its properties when it is completely isolated from other atoms.

40-2 Some Properties of Atoms

You may think the details of atomic physics are remote from your daily life. However, consider how the following properties of atoms—so basic that we rarely think about them—affect the way we live in our world.

Atoms are stable. Essentially all the atoms that form our tangible world have existed without change for billions of years. What would the world be like if atoms continually changed into other forms, perhaps every few weeks or every few years?

Atoms combine with each other. They stick together to form stable molecules and stack up to form rigid solids. An atom is mostly empty space, but you can stand on a floor—made up of atoms—without falling through it.

These basic properties of atoms can be explained by quantum physics, as can the three less apparent properties that follow.

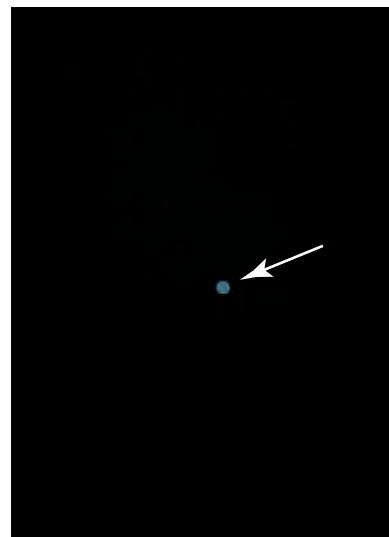


Fig. 40-1 The blue dot is a photograph of the light emitted from a single barium ion held for a long time in a trap at the University of Washington. Special techniques caused the ion to emit light over and over again as it underwent transitions between the same pair of energy levels. The dot represents the cumulative emission of many photons.
(Courtesy Warren Nagourney)

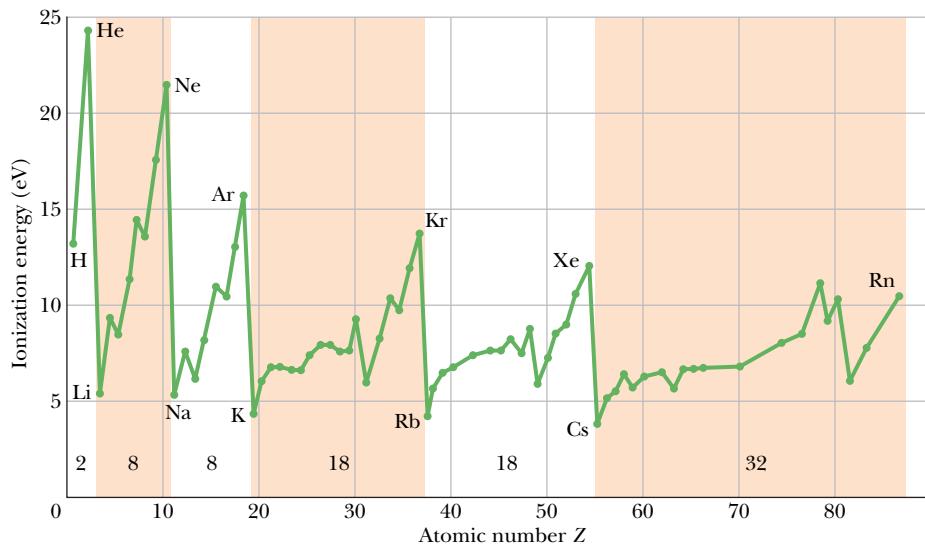


Fig. 40-2 A plot of the ionization energies of the elements as a function of atomic number, showing the periodic repetition of properties through the six complete horizontal periods of the periodic table. The number of elements in each of these periods is indicated.

Atoms Are Put Together Systematically

Figure 40-2 shows an example of a repetitive property of the elements as a function of their position in the periodic table (Appendix G). The figure is a plot of the **ionization energy** of the elements; the energy required to remove the most loosely bound electron from a neutral atom is plotted as a function of the position in the periodic table of the element to which the atom belongs. The remarkable similarities in the chemical and physical properties of the elements in each vertical column of the periodic table are evidence enough that the atoms are constructed according to systematic rules.

The elements are arranged in the periodic table in six complete horizontal **periods** (and a seventh incomplete period); except for the first, each period starts at the left with a highly reactive alkali metal (lithium, sodium, potassium, and so on) and ends at the right with a chemically inert noble gas (neon, argon, krypton, and so on). Quantum physics accounts for the chemical properties of these elements. The numbers of elements in the six periods are

$$2, 8, 8, 18, 18, \text{ and } 32.$$

Quantum physics predicts these numbers.

Atoms Emit and Absorb Light

We have already seen that atoms can exist only in discrete quantum states, each state having a certain energy. An atom can make a transition from one state to another by emitting light (to jump to a lower energy level E_{low}) or by absorbing light (to jump to a higher energy level E_{high}). As we first discussed in Section 39-3, the light is emitted or absorbed as a photon with energy

$$hf = E_{\text{high}} - E_{\text{low}}. \quad (40-1)$$

Thus, the problem of finding the frequencies of light emitted or absorbed by an atom reduces to the problem of finding the energies of the quantum states of that atom. Quantum physics allows us—in principle at least—to calculate these energies.

Atoms Have Angular Momentum and Magnetism

Figure 40-3 shows a negatively charged particle moving in a circular orbit around a fixed center. As we discussed in Section 32-7, the orbiting particle has both an

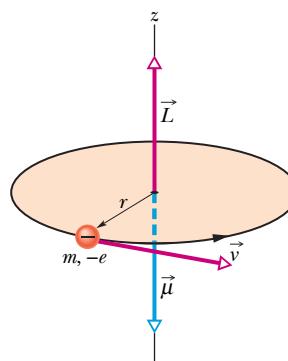


Fig. 40-3 A classical model showing a particle of mass m and charge $-e$ moving with speed v in a circle of radius r . The moving particle has an angular momentum \vec{L} given by $\vec{r} \times \vec{p}$, where \vec{p} is its linear momentum $mv\vec{v}$. The particle's motion is equivalent to a current loop that has an associated magnetic moment $\vec{\mu}$ that is directed opposite \vec{L} .

angular momentum \vec{L} and (because its path is equivalent to a tiny current loop) a magnetic dipole moment $\vec{\mu}$. As Fig. 40-3 shows, vectors \vec{L} and $\vec{\mu}$ are both perpendicular to the plane of the orbit but, because the charge is negative, they point in opposite directions.

The model of Fig. 40-3 is strictly classical and does not accurately represent an electron in an atom. In quantum physics, the rigid orbit model has been replaced by the probability density model, best visualized as a dot plot. In quantum physics, however, it is still true that in general, each quantum state of an electron in an atom involves an angular momentum \vec{L} and a magnetic dipole moment $\vec{\mu}$ that have opposite directions (those vector quantities are said to be *coupled*).

The Einstein-de Haas Experiment

In 1915, well before the discovery of quantum physics, Albert Einstein and Dutch physicist W. J. de Haas carried out a clever experiment designed to show that the angular momentum and magnetic moment of individual atoms are coupled.

Einstein and de Haas suspended an iron cylinder from a thin fiber, as shown in Fig. 40-4. A solenoid was placed around the cylinder but not touching it. Initially, the magnetic dipole moments $\vec{\mu}$ of the atoms of the cylinder point in random directions, and so their external magnetic effects cancel (Fig. 40-4a). However, when a current is switched on in the solenoid (Fig. 40-4b) so that a magnetic field \vec{B} is set up parallel to the long axis of the cylinder, the magnetic dipole moments of the atoms of the cylinder reorient themselves, lining up with that field. If the angular momentum \vec{L} of each atom is coupled to its magnetic moment $\vec{\mu}$, then this alignment of the atomic magnetic moments must cause an alignment of the atomic angular momenta opposite the magnetic field.

No external torques initially act on the cylinder; thus, its angular momentum must remain at its initial zero value. However, when \vec{B} is turned on and the atomic angular momenta line up antiparallel to \vec{B} , they tend to give a net angular momentum \vec{L}_{net} to the cylinder as a whole (directed downward in Fig. 40-4b). To maintain zero angular momentum, the cylinder begins to rotate around its central axis to produce an angular momentum \vec{L}_{rot} in the opposite direction (upward in Fig. 40-4b).

The twisting of the fiber quickly produces a torque that momentarily stops the cylinder's rotation and then rotates the cylinder in the opposite direction as the twisting is undone. Thereafter, the fiber will twist and untwist as the cylinder oscillates about its initial orientation in angular simple harmonic motion.

Observation of the cylinder's rotation verified that the angular momentum and the magnetic dipole moment of an atom are coupled in opposite directions. Moreover, it dramatically demonstrated that the angular momenta associated with quantum states of atoms can result in *visible* rotation of an object of everyday size.

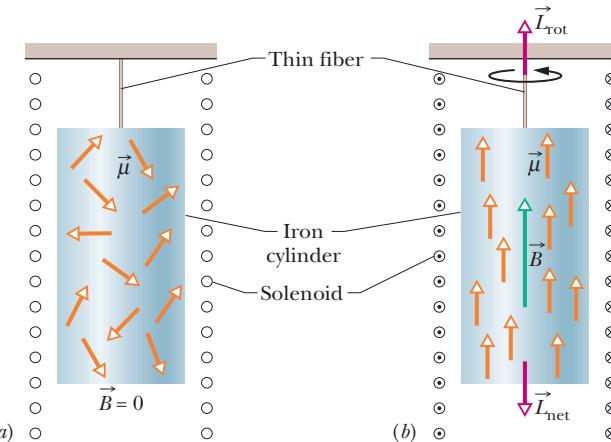


Fig. 40-4 The Einstein-de Haas experimental setup. (a) Initially, the magnetic field in the iron cylinder is zero and the magnetic dipole moment vectors $\vec{\mu}$ of its atoms are randomly oriented. (b) When a magnetic field \vec{B} is set up along the cylinder's axis, the magnetic dipole moment vectors line up parallel to \vec{B} and the cylinder begins to rotate.

Aligning the magnetic moment vectors rotates the cylinder.

40-3 Electron Spin

As we discussed in Section 32-7, whether an electron is *trapped* in an atom or is *free*, it has an intrinsic **spin angular momentum** \vec{S} , often called simply **spin**. (Recall that *intrinsic* means that \vec{S} is a basic characteristic of an electron, like its mass and electric charge.) As we shall discuss in the next section, the magnitude of \vec{S} is quantized and depends on a **spin quantum number** s , which is always $\frac{1}{2}$ for electrons (and for protons and neutrons). In addition, the component of \vec{S} measured along any axis is quantized and depends on a **spin magnetic quantum number** m_s , which can have only the value $+\frac{1}{2}$ or $-\frac{1}{2}$.

The existence of electron spin was postulated on an empirical basis by two Dutch graduate students, George Uhlenbeck and Samuel Goudsmit, from their studies of atomic spectra. The quantum physics basis for electron spin was provided a few years later, by British physicist P. A. M. Dirac, who developed (in 1929) a relativistic quantum theory of the electron.

It is tempting to account for electron spin by thinking of the electron as a tiny sphere spinning about an axis. However, that classical model, like the classical model of orbits, does not hold up. In quantum physics, spin angular momentum is best thought of as a measurable intrinsic property of the electron.

In Section 39-9, we briefly discussed the quantum numbers generated by applying Schrödinger's equation to the electron in a hydrogen atom (Table 39-2). We can now extend the list of quantum numbers by including s and m_s , as shown in Table 40-1. This set of five quantum numbers completely specifies the quantum state of an electron in a hydrogen atom or any other atom. All states with the same value of n form a **shell**. By counting the allowed values of ℓ and m_ℓ and then doubling the number to account for the two allowed values of m_s , you can verify that a shell defined by quantum number n has $2n^2$ states. All states with the same value of n and ℓ form a **subshell** and have the same energy. You can verify that a subshell defined by quantum number ℓ has $2(2\ell + 1)$ states.

40-4 Angular Momenta and Magnetic Dipole Moments

Every quantum state of an electron in an atom has an associated orbital angular momentum and a corresponding orbital magnetic dipole moment. Every electron, whether trapped in an atom or free, has a spin angular momentum and a corresponding spin magnetic dipole moment. Let's discuss these quantities.

Orbital Angular Momentum and Magnetism

The magnitude L of the **orbital angular momentum** \vec{L} of an electron *in an atom* is quantized; that is, it can have only certain values. These values are

$$L = \sqrt{\ell(\ell + 1)}\hbar, \quad (40-2)$$

Table 40-1

Electron States for an Atom

Quantum Number	Symbol	Allowed Values	Related to
Principal	n	1, 2, 3, ...	Distance from the nucleus
Orbital	ℓ	0, 1, 2, ..., $(n - 1)$	Orbital angular momentum
Orbital magnetic	m_ℓ	0, $\pm 1, \pm 2, \dots, \pm \ell$	Orbital angular momentum (z component)
Spin	s	$\frac{1}{2}$	Spin angular momentum
Spin magnetic	m_s	$\pm \frac{1}{2}$	Spin angular momentum (z component)

in which ℓ is the orbital quantum number and \hbar is $h/2\pi$. According to Table 40-1, ℓ must be either zero or a positive integer no greater than $n - 1$. For a state with $n = 3$, for example, only $\ell = 2$, $\ell = 1$, and $\ell = 0$ are permitted.

As we discussed in Section 32-7, a magnetic dipole is associated with the orbital angular momentum \vec{L} of an electron in an atom. This magnetic dipole has an **orbital magnetic dipole moment** $\vec{\mu}_{\text{orb}}$, which is related to the angular momentum by Eq. 32-28:

$$\vec{\mu}_{\text{orb}} = -\frac{e}{2m} \vec{L}. \quad (40-3)$$

The minus sign in this relation means that $\vec{\mu}_{\text{orb}}$ is directed opposite \vec{L} . Because the magnitude of \vec{L} is quantized (Eq. 40-2), the magnitude of $\vec{\mu}_{\text{orb}}$ must also be quantized and given by

$$\mu_{\text{orb}} = \frac{e}{2m} \sqrt{\ell(\ell + 1)}\hbar. \quad (40-4)$$

Neither $\vec{\mu}_{\text{orb}}$ nor \vec{L} can be measured in any way. However, we *can* measure their components along a given axis. Suppose that the atom is located in a magnetic field \vec{B} , with a z axis extending in the direction of the field lines at the atom's location. Then we can measure the z components of $\vec{\mu}_{\text{orb}}$ and \vec{L} along that axis.

The components $\mu_{\text{orb},z}$ are quantized and given by

$$\mu_{\text{orb},z} = -m_\ell \mu_B. \quad (40-5)$$

Here m_ℓ is the orbital magnetic quantum number of Table 40-1 and μ_B is the **Bohr magneton**:

$$\mu_B = \frac{e\hbar}{4\pi m} = \frac{e\hbar}{2m} = 9.274 \times 10^{-24} \text{ J/T} \quad (\text{Bohr magneton}), \quad (40-6)$$

where m is the electron mass.

The components L_z of the angular momentum are also quantized, and they are given by

$$L_z = m_\ell \hbar. \quad (40-7)$$

Figure 40-5 shows the five quantized components L_z of the orbital angular momentum for an electron with $\ell = 2$, as well as the associated orientations of the angular momentum \vec{L} . However, *do not take the figure literally* because we cannot detect \vec{L} in any way. Thus, drawing it in a figure like Fig. 40-5 is merely a visual aide. We can extend that visual aide by saying that \vec{L} makes a certain angle θ with the z axis, such that

$$\cos \theta = \frac{L_z}{L}. \quad (40-8)$$

We can call θ the *semi-classical angle* between vector \vec{L} and the z axis because θ is a classical measure of something that quantum theory tells us cannot be measured.

Spin Angular Momentum and Spin Magnetic Dipole Moment

The magnitude S of the spin angular momentum \vec{S} of any electron, whether *free or trapped*, has the single value given by

$$\begin{aligned} S &= \sqrt{s(s + 1)}\hbar \\ &= \sqrt{(\frac{1}{2})(\frac{1}{2} + 1)}\hbar = 0.866\hbar, \end{aligned} \quad (40-9)$$

where $s (= \frac{1}{2})$ is the spin quantum number of the electron.

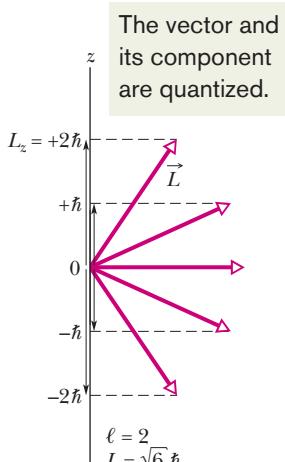


Fig. 40-5 The allowed values of L_z for an electron in a quantum state with $\ell = 2$. For every orbital angular momentum vector \vec{L} in the figure, there is a vector pointing in the opposite direction, representing the magnitude and direction of the orbital magnetic dipole moment $\vec{\mu}_{\text{orb}}$.

40-4 ANGULAR MOMENTA AND MAGNETIC DIPOLE MOMENTS

1117

As we discussed in Section 32-7, an electron has an intrinsic magnetic dipole that is associated with its spin angular momentum \vec{S} , whether the electron is confined to an atom or free. This magnetic dipole has a **spin magnetic dipole moment** $\vec{\mu}_s$, which is related to the spin angular momentum \vec{S} by Eq. 32-22:

$$\vec{\mu}_s = -\frac{e}{m} \vec{S}. \quad (40-10)$$

The minus sign in this relation means that $\vec{\mu}_s$ is directed opposite \vec{S} . Because the magnitude of \vec{S} is quantized (Eq. 40-9), the magnitude of $\vec{\mu}_s$ must also be quantized and given by

$$\mu_s = \frac{e}{m} \sqrt{s(s+1)}\hbar. \quad (40-11)$$

Neither \vec{S} nor $\vec{\mu}_s$ can be measured in any way. However, we *can* measure their components along any given axis—call it the z axis. The components S_z of the spin angular momentum are quantized and given by

$$S_z = m_s \hbar, \quad (40-12)$$

where m_s is the spin magnetic quantum number of Table 40-1. That quantum number can have only two values: $m_s = +\frac{1}{2}$ (the electron is said to be *spin up*) and $m_s = -\frac{1}{2}$ (the electron is said to be *spin down*).

The components $\mu_{s,z}$ of the spin magnetic dipole moment are also quantized, and they are given by

$$\mu_{s,z} = -2m_s \mu_B. \quad (40-13)$$

Figure 40-6 shows the two quantized components S_z of the spin angular momentum for an electron and the associated orientations of vector \vec{S} . It also shows the quantized components $\mu_{s,z}$ of the spin magnetic dipole moment and the associated orientations of $\vec{\mu}_s$.

Orbital and Spin Angular Momenta Combined

For an atom containing more than one electron, we define a total angular momentum \vec{J} , which is the vector sum of the angular momenta of the individual electrons—both their orbital and their spin angular momenta. Each element in the periodic table is defined by the number of protons in the nucleus of an atom of the element. This number of protons is defined as being the **atomic number** (or **charge number**) Z of the element. Because an electrically neutral atom contains equal numbers of protons and electrons, Z is also the number of electrons in the neutral atom, and we use this fact to indicate a \vec{J} value for a neutral atom:

$$\vec{J} = (\vec{L}_1 + \vec{L}_2 + \vec{L}_3 + \cdots + \vec{L}_Z) + (\vec{S}_1 + \vec{S}_2 + \vec{S}_3 + \cdots + \vec{S}_Z). \quad (40-14)$$

Similarly, the total magnetic dipole moment of a multielectron atom is the vector sum of the magnetic dipole moments (both orbital and spin) of its individual electrons. However, because of the factor 2 in Eq. 40-13, the resultant magnetic dipole moment for the atom does not have the direction of vector $-\vec{J}$; instead, it makes a certain angle with that vector. The **effective magnetic dipole moment** $\vec{\mu}_{\text{eff}}$ for the atom is the component of the vector sum of the individual magnetic dipole moments in the direction of $-\vec{J}$ (Fig. 40-7). In typical atoms the orbital angular momenta and the spin angular momenta of most of the electrons sum vectorially to zero. Then \vec{J} and $\vec{\mu}_{\text{eff}}$ of those atoms are due to a relatively small number of electrons, often only a single valence electron.

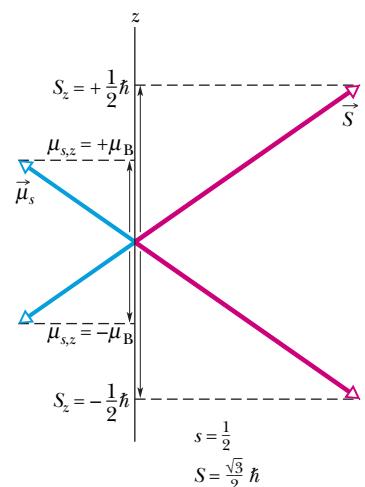


Fig. 40-6 The allowed values of S_z and μ_z for an electron.

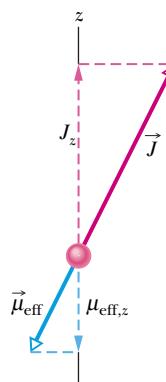


Fig. 40-7 A classical model showing the total angular momentum vector \vec{J} and the effective magnetic moment vector $\vec{\mu}_{\text{eff}}$.



CHECKPOINT 1

An electron is in a quantum state for which the magnitude of the electron's orbital angular momentum \vec{L} is $2\sqrt{3}\hbar$. How many projections of the electron's orbital magnetic dipole moment on a z axis are allowed?

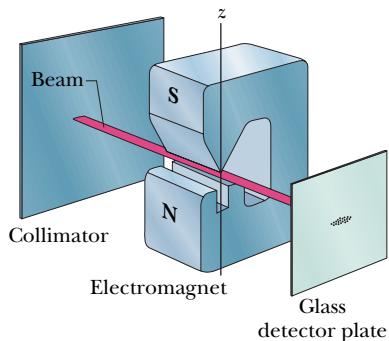


Fig. 40-8 Apparatus used by Stern and Gerlach.

40-5 The Stern-Gerlach Experiment

In 1922, Otto Stern and Walther Gerlach at the University of Hamburg in Germany showed experimentally that the magnetic moment of silver atoms is quantized. In the Stern–Gerlach experiment, as it is now known, silver is vaporized in an oven, and some of the atoms in that vapor escape through a narrow slit in the oven wall and pass into an evacuated tube. Some of those escaping atoms then pass through a second narrow slit, to form a narrow beam of atoms (Fig. 40-8). (The atoms are said to be *collimated*—made into a beam—and the second slit is called a *collimator*.) The beam passes between the poles of an electromagnet and then lands on a glass detector plate where it forms a silver deposit.

When the electromagnet is off, the silver deposit is a narrow spot. However, when the electromagnet is turned on, the silver deposit should be spread vertically. The reason is that silver atoms are magnetic dipoles, and so vertical magnetic forces act on them as they pass through the vertical magnetic field of the electromagnet; these forces deflect them slightly up or down. Thus, by analyzing the silver deposit on the plate, we can determine what deflections the atoms underwent in the magnetic field. When Stern and Gerlach analyzed the pattern of silver on their detector plate, they found a surprise. However, before we discuss that surprise and its quantum implications, let us discuss the magnetic deflecting force acting on the silver atoms.

The Magnetic Deflecting Force on a Silver Atom

We have not previously discussed the type of magnetic force that deflects the silver atoms in a Stern–Gerlach experiment. It is *not* the magnetic deflecting force that acts on a moving charged particle, as given by Eq. 28-2 ($\vec{F} = q\vec{v} \times \vec{B}$). The reason is simple: A silver atom is electrically neutral (its net charge q is zero), and thus this type of magnetic force is also zero.

The type of magnetic force we seek is due to an interaction between the magnetic field \vec{B} of the electromagnet and the magnetic dipole of the individual silver atom. We can derive an expression for the force in this interaction by starting with the energy U of the dipole in the magnetic field. Equation 28-38 tells us that

$$U = -\vec{\mu} \cdot \vec{B}, \quad (40-15)$$

where $\vec{\mu}$ is the magnetic dipole moment of a silver atom. In Fig. 40-8, the positive direction of the z axis and the direction of \vec{B} are vertically upward. Thus, we can write Eq. 40-15 in terms of the component μ_z of the atom's magnetic dipole moment along the direction of \vec{B} :

$$U = -\mu_z B. \quad (40-16)$$

Then, using Eq. 8-22 ($F = -dU/dx$) for the z axis shown in Fig. 40-8, we obtain

$$F_z = -\frac{dU}{dz} = \mu_z \frac{dB}{dz}. \quad (40-17)$$

This is what we sought—an equation for the magnetic force that deflects a silver atom as the atom passes through a magnetic field.

The term dB/dz in Eq. 40-17 is the *gradient* of the magnetic field along the z axis. If the magnetic field does not change along the z axis (as in a uniform magnetic field or no magnetic field), then $dB/dz = 0$ and a silver atom is not deflected as it moves between the magnet's poles. In the Stern–Gerlach experiment, the poles are designed to maximize the gradient dB/dz , so as to vertically deflect the silver atoms passing between the poles as much as possible, so that their deflections show up in the deposit on the glass plate.

According to classical physics, the components μ_z of silver atoms passing through the magnetic field in Fig. 40-8 should range in value from $-\mu$ (the dipole

40-5 THE STERN-GERLACH EXPERIMENT

1119

moment $\vec{\mu}$ is directed straight down the z axis) to $+ \mu$ ($\vec{\mu}$ is directed straight up the z axis). Thus, from Eq. 40-17, there should be a range of forces on the atoms, and therefore a range of deflections of the atoms, from a greatest downward deflection to a greatest upward deflection. This means that we should expect the atoms to land along a vertical line on the glass plate, but they *don't*.

The Experimental Surprise

What Stern and Gerlach found was that the atoms formed two distinct spots on the glass plate, one spot above the point where they would have landed with no deflection and the other spot just as far below that point. The spots were initially too faint to be seen, but they became visible when Stern happened to breathe on the glass plate after smoking a cheap cigar. Sulfur in his breath (from the cigar) combined with the silver to produce a noticeably black silver sulfide.

This two-spot result can be seen in the plots of Fig. 40-9, which shows the outcome of a more recent version of the Stern–Gerlach experiment. In that version, a beam of cesium atoms (magnetic dipoles like the silver atoms in the original Stern–Gerlach experiment) was sent through a magnetic field with a large vertical gradient dB/dz . The field could be turned on and off, and a detector could be moved up and down through the beam.

When the field was turned off, the beam was, of course, undeflected and the detector recorded the central-peak pattern shown in Fig. 40-9. When the field was turned on, the original beam was split vertically by the magnetic field into two smaller beams, one beam higher than the previously undeflected beam and the other beam lower. As the detector moved vertically up through these two smaller beams, it recorded the two-peak pattern shown in Fig. 40-9.

The Meaning of the Results

In the original Stern–Gerlach experiment, two spots of silver were formed on the glass plate, not a vertical line of silver. This means that the component μ_z along \vec{B} (and along z) could not have any value between $-\mu$ and $+\mu$ as classical physics predicts. Instead, μ_z is restricted to only two values, one for each spot on the glass. Thus, the original Stern–Gerlach experiment showed that μ_z is quantized, implying (correctly) that $\vec{\mu}$ is also. Moreover, because the angular momentum \vec{L} of an atom is associated with $\vec{\mu}$, that angular momentum and its component L_z are also quantized.

With modern quantum theory, we can add to the explanation of the two-spot result in the Stern–Gerlach experiment. We now know that a silver atom consists of many electrons, each with a spin magnetic moment and an orbital magnetic moment. We also know that all those moments vectorially cancel out *except* for a single electron, and the orbital dipole moment of that electron is zero. Thus, the combined dipole moment $\vec{\mu}$ of a silver atom is the *spin* magnetic dipole moment of that single electron. According to Eq. 40-13, this means that μ_z can have only two components along the z axis in Fig. 40-8. One component is for quantum number $m_s = +\frac{1}{2}$ (the single electron is spin up), and the other component is for quantum number $m_s = -\frac{1}{2}$ (the single electron is spin down). Substituting into Eq. 40-13 gives us

$$\mu_{s,z} = -2(+\frac{1}{2})\mu_B = -\mu_B \quad \text{and} \quad \mu_{s,z} = -2(-\frac{1}{2})\mu_B = +\mu_B. \quad (40-18)$$

Then substituting these expressions for μ_z in Eq. 40-17, we find that the force component F_z deflecting the silver atoms as they pass through the magnetic field can have only the two values

$$F_z = -\mu_B \left(\frac{dB}{dz} \right) \quad \text{and} \quad F_z = +\mu_B \left(\frac{dB}{dz} \right), \quad (40-19)$$

which result in the two spots of silver on the glass.

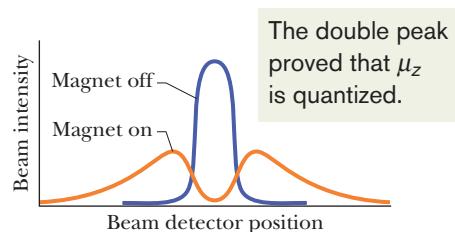


Fig. 40-9 Results of a modern repetition of the Stern–Gerlach experiment. With the electromagnet turned off, there is only a single beam; with the electromagnet turned on, the original beam splits into two subbeams. The two subbeams correspond to parallel and antiparallel alignment of the magnetic moments of cesium atoms with the external magnetic field.

Sample Problem**Beam separation in a Stern–Gerlach experiment**

In the Stern–Gerlach experiment of Fig. 40-8, a beam of silver atoms passes through a magnetic field gradient dB/dz of magnitude 1.4 T/mm that is set up along the z axis. This region has a length w of 3.5 cm in the direction of the original beam. The speed of the atoms is 750 m/s. By what distance d have the atoms been deflected when they leave the region of the field gradient? The mass M of a silver atom is 1.8×10^{-25} kg.

KEY IDEAS

(1) The deflection of a silver atom in the beam is due to an interaction between the magnetic dipole of the atom and the magnetic field, because of the gradient dB/dz . The deflecting force is directed along the field gradient (along the z axis) and is given by Eqs. 40-19. Let us consider only deflection in the positive direction of z ; thus, we shall use $F_z = \mu_B(dB/dz)$ from Eqs. 40-19.

(2) We assume the field gradient dB/dz has the same value throughout the region through which the silver atoms travel. Thus, force component F_z is constant in that region, and from Newton's second law, the acceleration a_z of an atom along the z axis due to F_z is also constant.

Calculations: Putting these ideas together, we write the acceleration as

$$a_z = \frac{F_z}{M} = \frac{\mu_B(dB/dz)}{M}.$$

Because this acceleration is constant, we can use Eq. 2-15 (from Table 2-1) to write the deflection d parallel to the z axis as

$$d = v_{0z}t + \frac{1}{2}a_z t^2 = 0t + \frac{1}{2}\left(\frac{\mu_B(dB/dz)}{M}\right)t^2. \quad (40-20)$$

Because the deflecting force on the atom acts perpendicular to the atom's original direction of travel, the component v of the atom's velocity along the original direction of travel is not changed by the force. Thus, the atom requires time $t = w/v$ to travel through length w in that direction. Substituting w/v for t into Eq. 40-20, we find

$$\begin{aligned} d &= \frac{1}{2}\left(\frac{\mu_B(dB/dz)}{M}\right)\left(\frac{w}{v}\right)^2 = \frac{\mu_B(dB/dz)w^2}{2Mv^2} \\ &= (9.27 \times 10^{-24} \text{ J/T})(1.4 \times 10^3 \text{ T/m}) \\ &\quad \times \frac{(3.5 \times 10^{-2} \text{ m})^2}{(2)(1.8 \times 10^{-25} \text{ kg})(750 \text{ m/s})^2} \\ &= 7.85 \times 10^{-5} \text{ m} \approx 0.08 \text{ mm}. \end{aligned} \quad (\text{Answer})$$

The separation between the two subbeams is twice this, or 0.16 mm. This separation is not large but is easily measured.



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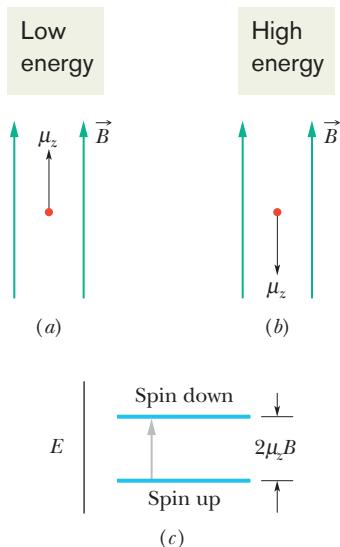


Fig. 40-10 The z component of $\vec{\mu}$ for a proton in the (a) lower-energy (spin-up) and (b) higher-energy (spin-down) state. (c) An energy-level diagram for the states, showing the upward quantum jump the proton makes when its spin flips from up to down.

40-6 Magnetic Resonance

As we discussed briefly in Section 32-7, a proton has a spin magnetic dipole moment $\vec{\mu}$ that is associated with the proton's intrinsic spin angular momentum \vec{S} . The two vectors are said to be coupled together and, because the proton is positively charged, they are in the same direction. Suppose a proton is located in a magnetic field \vec{B} that is directed along the positive direction of a z axis. Then $\vec{\mu}$ has two possible quantized components along that axis: the component can be $+\mu_z$ if the vector is in the direction of \vec{B} (Fig. 40-10a) or $-\mu_z$ if it is opposite the direction of \vec{B} (Fig. 40-10b).

From Eq. 28-38 ($U(\theta) = -\vec{\mu} \cdot \vec{B}$), recall that an energy is associated with the orientation of any magnetic dipole moment $\vec{\mu}$ located in an external magnetic field \vec{B} . Thus, energy is associated with the two orientations of Figs. 40-10a and b. The orientation in Fig. 40-10a is the lower-energy state ($-\mu_z B$) and is called the *spin-up state* because the proton's spin component S_z (not shown) is also aligned with \vec{B} . The orientation in Fig. 40-10b (the *spin-down state*) is the higher-energy state ($\mu_z B$). Thus, the energy difference between these two states is

$$\Delta E = \mu_z B - (-\mu_z B) = 2\mu_z B. \quad (40-21)$$

If we place a sample of water in a magnetic field \vec{B} , the protons in the hydrogen portions of each water molecule tend to be in the lower-energy state. (We shall not consider the oxygen portions.) Any one of these protons can jump to the higher-energy state by absorbing a photon with an energy hf equal to ΔE . That is,

40-8 MULTIPLE ELECTRONS IN RECTANGULAR TRAPS

1121

the proton can jump by absorbing a photon of energy

$$hf = 2\mu_z B. \quad (40-22)$$

Such absorption is called **magnetic resonance** or, as originally, **nuclear magnetic resonance** (NMR), and the consequent reversal of S_z is called *spin-flipping*.

In practice, the photons required for magnetic resonance have an associated frequency in the radio-frequency (RF) range and are provided by a small coil wrapped around the sample undergoing resonance. An electromagnetic oscillator called an *RF source* drives a sinusoidal current in the coil at frequency f . The electromagnetic (EM) field set up within the coil and sample also oscillates at frequency f . If f meets the requirement of Eq. 40-22, the oscillating EM field can transfer a quantum of energy to a proton in the sample via a photon absorption, spin-flipping the proton.

The magnetic field magnitude B that appears in Eq. 40-22 is actually the magnitude of the net magnetic field \vec{B} at the site where a given proton undergoes spin-flipping. That net field is the vector sum of the external field \vec{B}_{ext} set up by the magnetic resonance equipment (primarily a large magnet) and the internal field \vec{B}_{int} set up by the magnetic dipole moments of the atoms and nuclei near the given proton. For practical reasons we do not discuss here, magnetic resonance is usually detected by sweeping the magnitude B_{ext} through a range of values while the frequency f of the RF source is kept at a predetermined value and the energy of the RF source is monitored. A graph of the energy loss of the RF source versus B_{ext} shows a *resonance peak* when B_{ext} sweeps through the value at which spin-flipping occurs. Such a graph is called a *nuclear magnetic resonance spectrum*, or *NMR spectrum*.

Figure 40-11 shows the NMR spectrum of ethanol, which is a molecule consisting of three groups of atoms: CH_3 , CH_2 , and OH. Protons in each group can undergo magnetic resonance, but each group has its own unique magnetic-resonance value of B_{ext} because the groups lie in different internal fields \vec{B}_{int} due to their arrangement within the $\text{CH}_3\text{CH}_2\text{OH}$ molecule. Thus, the resonance peaks in the spectrum of Fig. 40-11 form a unique NMR signature by which ethanol can be identified.

40-7 The Pauli Exclusion Principle

In Chapter 39 we considered a variety of electron traps, from fictional one-dimensional traps to the real three-dimensional trap of a hydrogen atom. In all those examples, we trapped only one electron. However, when we discuss traps containing two or more electrons (as we shall in the next two sections), we must consider a principle that governs any particle whose spin quantum number s is not zero or an integer. This principle applies not only to electrons but also to protons and neutrons, all of which have $s = \frac{1}{2}$. The principle is known as the **Pauli exclusion principle** after Wolfgang Pauli, who formulated it in 1925. For electrons, it states that



No two electrons confined to the same trap can have the same set of values for their quantum numbers.

As we shall discuss in Section 40-9, this principle means that no two electrons in an atom can have the same four values for the quantum numbers n , ℓ , m_ℓ , and m_s . All electrons have the same quantum number $s = \frac{1}{2}$. Thus, any two electrons in an atom must differ in at least one of these other quantum numbers. Were this not true, atoms would collapse, and thus you and the world could not exist.

40-8 Multiple Electrons in Rectangular Traps

To prepare for our discussion of multiple electrons in atoms, let us discuss two electrons confined to the rectangular traps of Chapter 39. We shall again use the

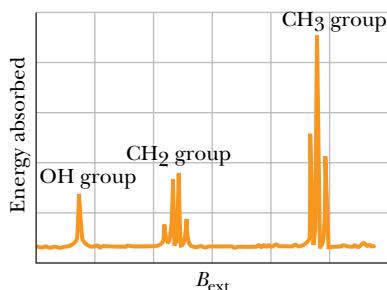


Fig. 40-11 A nuclear magnetic resonance spectrum for ethanol, $\text{CH}_3\text{CH}_2\text{OH}$. The spectral lines represent the absorption of energy associated with spin-flips of protons. The three groups of lines correspond, as indicated, to protons in the OH group, the CH_2 group, and the CH_3 group of the ethanol molecule. Note that the two protons in the CH_2 group occupy four different local environments. The entire horizontal axis covers less than 10^{-4} T.

quantum numbers we found for those traps when only one electron was confined. However, here we shall also include the spin angular momenta of the two electrons. To do this, we assume that the traps are located in a uniform magnetic field. Then according to Eq. 40-12 ($S_z = m_s \hbar$), an electron can be either spin up with $m_s = \frac{1}{2}$ or spin down with $m_s = -\frac{1}{2}$. (We assume that the field is very weak so that we can neglect the energies of the electrons due to it.)

As we confine the two electrons to one of the traps, we must keep the Pauli exclusion principle in mind; that is, the electrons cannot have the same set of values for their quantum numbers.

1. *One-dimensional trap.* In the one-dimensional trap of Fig. 39-2, fitting an electron wave to the trap's width L requires the single quantum number n . Therefore, any electron confined to the trap must have a certain value of n , and its quantum number m_s can be either $+\frac{1}{2}$ or $-\frac{1}{2}$. The two electrons could have different values of n , or they could have the same value of n if one of them is spin up and the other is spin down.
2. *Rectangular corral.* In the rectangular corral of Fig. 39-13, fitting an electron wave to the corral's widths L_x and L_y requires the two quantum numbers n_x and n_y . Thus, any electron confined to the trap must have certain values for those two quantum numbers, and its quantum number m_s can be either $+\frac{1}{2}$ or $-\frac{1}{2}$; so now there are three quantum numbers. According to the Pauli exclusion principle, two electrons confined to the trap must have different values for at least one of those three quantum numbers.
3. *Rectangular box.* In the rectangular box of Fig. 39-14, fitting an electron wave to the box's widths L_x , L_y , and L_z requires the three quantum numbers n_x , n_y , and n_z . Thus, any electron confined to the trap must have certain values for these three quantum numbers, and its quantum number m_s can be either $+\frac{1}{2}$ or $-\frac{1}{2}$; so now there are four quantum numbers. According to the Pauli exclusion principle, two electrons confined to the trap must have different values for at least one of those four quantum numbers.

Suppose we add more than two electrons, one by one, to a rectangular trap in the preceding list. The first electrons naturally go into the lowest possible energy level—they are said to *occupy* that level. However, eventually the Pauli exclusion principle disallows any more electrons from occupying that lowest energy level, and the next electron must occupy the next higher level. When an energy level cannot be occupied by more electrons because of the Pauli exclusion principle, we say that level is **full** or **fully occupied**. In contrast, a level that is not occupied by any electrons is **empty** or **unoccupied**. For intermediate situations, the level is **partially occupied**. The *electron configuration* of a system of trapped electrons is a listing or drawing either of the energy levels the electrons occupy or of the set of the quantum numbers of the electrons.

Finding the Total Energy

To find the energy of a system of two or more electrons confined to a trap, we assume that the electrons do not electrically interact with one another; that is, we shall neglect the electric potential energies of pairs of electrons. Then we can calculate the total energy for the system by calculating the energy of each electron (as in Chapter 39) and then summing those energies.

A good way to organize the energy values of a given system of electrons is with an energy-level diagram *for the system*, just as we did for a single electron in the traps of Chapter 39. The lowest level, with energy E_{gr} , corresponds to the ground state of the system. The next higher level, with energy E_{fe} , corresponds to the first excited state of the system. The next level, with energy E_{se} , corresponds to the second excited state of the system, and so on.

Sample Problem

Energy levels of multiple electrons in a 2D infinite potential well

Seven electrons are confined to a square corral (two-dimensional infinite potential well) with widths $L_x = L_y = L$ (Fig. 39-13). Assume that the electrons do not electrically interact with one another.

- (a) What is the electron configuration for the ground state of the system of seven electrons?

One-electron diagram: We can determine the electron configuration of the system by placing the seven electrons in the corral one by one, to build up the system. Because we assume the electrons do not electrically interact with one another, we can use the energy-level diagram for a single trapped electron in order to keep track of how we place the seven electrons in the corral. That *one-electron energy-level diagram* is given in Fig. 39-15 and partially reproduced here as Fig. 40-12a. Recall that the levels are labeled as $E_{nx,ny}$ for their associated energy. For example, the lowest level is for energy $E_{1,1}$, where quantum number n_x is 1 and quantum number n_y is 1.

Pauli principle: The trapped electrons must obey the Pauli exclusion principle; that is, no two electrons can have the same set of values for their quantum numbers n_x , n_y , and m_s . The first electron goes into energy level $E_{1,1}$ and can have $m_s = \frac{1}{2}$ or $m_s = -\frac{1}{2}$. We arbitrarily choose the lat-

ter and draw a down arrow (to represent spin down) on the $E_{1,1}$ level in Fig. 40-12a. The second electron also goes into the $E_{1,1}$ level but must have $m_s = +\frac{1}{2}$ so that one of its quantum numbers differs from those of the first electron. We represent this second electron with an up arrow (for spin up) on the $E_{1,1}$ level in Fig. 40-12b.

Electrons, one by one: The level for energy $E_{1,1}$ is fully occupied, and thus the third electron cannot have that energy. Therefore, the third electron goes into the next higher level, which is for the equal energies $E_{2,1}$ and $E_{1,2}$ (the level is degenerate). This third electron can have quantum numbers n_x and n_y of either 1 and 2 or 2 and 1, respectively. It can also have a quantum number m_s of either $+\frac{1}{2}$ or $-\frac{1}{2}$. Let us arbitrarily assign it the quantum numbers $n_x = 2$, $n_y = 1$, and $m_s = -\frac{1}{2}$. We then represent it with a down arrow on the level for $E_{1,2}$ and $E_{2,1}$ in Fig. 40-12c.

You can show that the next three electrons can also go into the level for energies $E_{2,1}$ and $E_{1,2}$, provided that no set of three quantum numbers is completely duplicated. That level then contains four electrons (Fig. 40-12d), with quantum numbers (n_s, n_v, m_s) of

$$(2, 1, -\tfrac{1}{2}), (2, 1, +\tfrac{1}{2}), (1, 2, -\tfrac{1}{2}), (1, 2, +\tfrac{1}{2}),$$

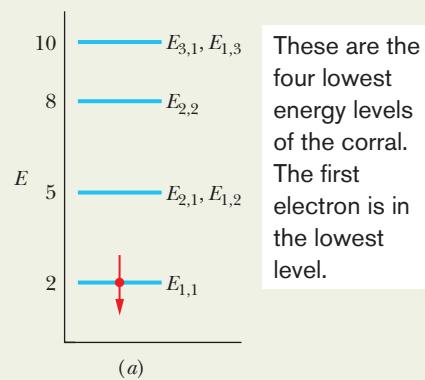
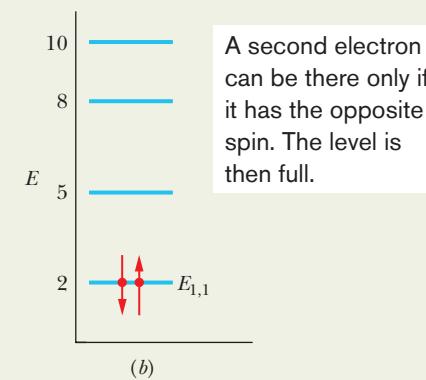
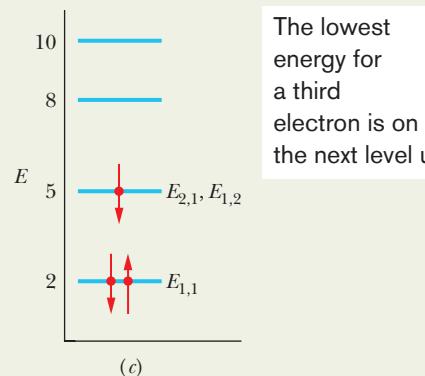


Fig. 40-12 (a) Energy-level diagram for one electron in a square corral. (Energy E is in multiples of $\hbar^2/8mL^2$.) A spin-down electron occupies the lowest level. (b) Two electrons (one spin down, the other spin up) occupy the lowest level of the one-electron energy-level diagram. (c) A third electron occupies the next energy level. (d) Four electrons can be put into the second level. (Figure continues on page 1125.)



The lowest energy for third electron is on the next level up.

Two quantum states have that energy. Two electrons (with opposite spins) can be in each state. Then that level is also full.

and the level is fully occupied. Thus, the seventh electron goes into the next higher level, which is the $E_{2,2}$ level. Let us assume this electron is spin down, with $m_s = -\frac{1}{2}$.

Figure 40-12e shows all seven electrons on a one-electron energy-level diagram. We now have seven electrons in the corral, and they are in the configuration with the lowest energy that satisfies the Pauli exclusion principle. Thus, the ground-state configuration of the system is that shown in Fig. 40-12e and listed in Table 40-2.

- (b) What is the total energy of the seven-electron system in its ground state, as a multiple of $\hbar^2/8mL^2$?

KEY IDEA

The total energy E_{gr} is the sum of the energies of the individual electrons in the system's ground-state configuration.

Ground-state energy: The energy of each electron can be read from Table 39-1, which is partially reproduced in Table 40-2, or from Fig. 40-12e. Because there are two electrons in the first (lowest) level, four in the second level, and one in the third level, we have

$$\begin{aligned} E_{\text{gr}} &= 2\left(2 \frac{\hbar^2}{8mL^2}\right) + 4\left(5 \frac{\hbar^2}{8mL^2}\right) + 1\left(8 \frac{\hbar^2}{8mL^2}\right) \\ &= 32 \frac{\hbar^2}{8mL^2}. \end{aligned} \quad (\text{Answer})$$

- (c) How much energy must be transferred to the system for it to jump to its first excited state, and what is the energy of that state?

KEY IDEAS

1. If the system is to be excited, one of the seven electrons must make a quantum jump up the one-electron energy-level diagram of Fig. 40-12e.

Table 40-2
Ground-State Configuration and Energies

n_x	n_y	m_s	Energy ^a
2	2	$-\frac{1}{2}$	8
2	1	$+\frac{1}{2}$	5
2	1	$-\frac{1}{2}$	5
1	2	$+\frac{1}{2}$	5
1	2	$-\frac{1}{2}$	5
1	1	$+\frac{1}{2}$	2
1	1	$-\frac{1}{2}$	$\frac{2}{2}$
			Total 32

^aIn multiples of $\hbar^2/8mL^2$.

2. If that jump is to occur, the energy change ΔE of the electron (and thus of the system) must be $\Delta E = E_{\text{high}} - E_{\text{low}}$ (Eq. 39-5), where E_{low} is the energy of the level where the jump begins and E_{high} is the energy of the level where the jump ends.
3. The Pauli exclusion principle must still apply; an electron *cannot* jump to a level that is fully occupied.

First-excited-state energy: Let us consider the three jumps shown in Fig. 40-12f; all are allowed by the Pauli exclusion principle because they are jumps to either empty or partially occupied states. In one of those possible jumps, an electron jumps from the $E_{1,1}$ level to the partially occupied $E_{2,2}$ level. The change in the energy is

$$\Delta E = E_{2,2} - E_{1,1} = 8 \frac{\hbar^2}{8mL^2} - 2 \frac{\hbar^2}{8mL^2} = 6 \frac{\hbar^2}{8mL^2}.$$

(We shall assume that the spin orientation of the electron making the jump can change as needed.)

40-9 Building the Periodic Table

The four quantum numbers n , ℓ , m_ℓ , and m_s identify the quantum states of individual electrons in a multielectron atom. The wave functions for these states, however, are not the same as the wave functions for the corresponding states of the hydrogen atom because, in multielectron atoms, the potential energy associated with a given electron is determined not only by the charge and position of the atom's nucleus but also by the charges and positions of all the other electrons in the atom. Solutions of Schrödinger's equation for multielectron atoms can be carried out numerically—in principle at least—using a computer.

As we discussed in Sections 39-9 and 40-3, all states with the same values of the quantum numbers n and ℓ form a subshell. For a given value of ℓ , there are $2\ell + 1$ possible values of the quantum

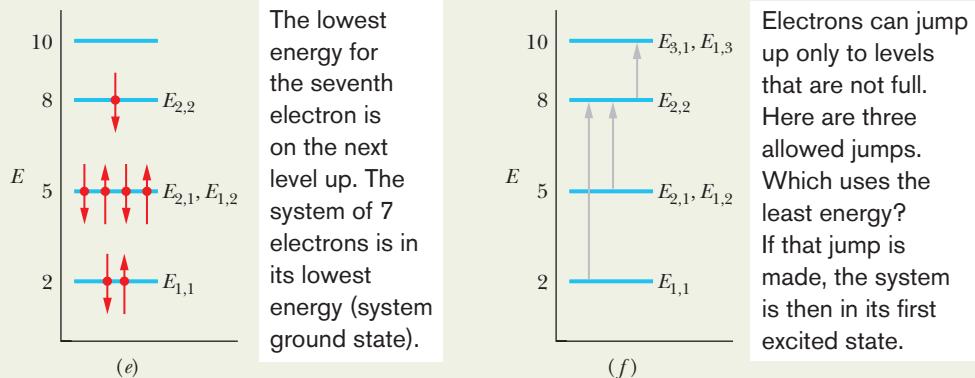


Fig. 40-12 (Continued from page 1123) (e) The system's ground-state configuration. (f) Three transitions to consider for the first excited state. (g) The system's lowest three total energies.

In another of the possible jumps in Fig. 40-12f, an electron jumps from the degenerate level of $E_{2,1}$ and $E_{1,2}$ to the partially occupied $E_{2,2}$ level. The change in the energy is

$$\Delta E = E_{2,2} - E_{2,1} = 8 \frac{h^2}{8mL^2} - 5 \frac{h^2}{8mL^2} = 3 \frac{h^2}{8mL^2}.$$

In the third possible jump in Fig. 40-12f, the electron in the $E_{2,2}$ level jumps to the unoccupied, degenerate level of $E_{1,3}$ and $E_{3,1}$. The change in energy is

$$\Delta E = E_{1,3} - E_{2,2} = 10 \frac{h^2}{8mL^2} - 8 \frac{h^2}{8mL^2} = 2 \frac{h^2}{8mL^2}.$$

Of these three possible jumps, the one requiring the least energy change ΔE is the last one. We could consider even more possible jumps, but none would require less energy.

Thus, for the system to jump from its ground state to its first excited state, the electron in the $E_{2,2}$ level must jump to the unoccupied, degenerate level of $E_{1,3}$ and $E_{3,1}$, and the required energy is

$$\Delta E = 2 \frac{h^2}{8mL^2}. \quad (\text{Answer})$$

The energy E_{fe} of the first excited state of the system is then

$$\begin{aligned} E_{\text{fe}} &= E_{\text{gr}} + \Delta E \\ &= 32 \frac{h^2}{8mL^2} + 2 \frac{h^2}{8mL^2} = 34 \frac{h^2}{8mL^2}. \quad (\text{Answer}) \end{aligned}$$

We can represent this energy and the energy E_{gr} for the ground state of the system on an energy-level diagram for the system, as shown in Fig. 40-12g.



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number m_ℓ and, for each m_ℓ , there are two possible values for the quantum number m_s . Thus, there are $2(2\ell + 1)$ states in a subshell. It turns out that *all states in a given subshell have the same energy*, its value being determined primarily by the value of n and to a lesser extent by the value of ℓ .

For the purpose of labeling subshells, the values of ℓ are represented by letters:

$$\begin{array}{ccccccc} \ell = 0 & 1 & 2 & 3 & 4 & 5 & \dots \\ s & p & d & f & g & h & \dots \end{array}$$

For example, the $n = 3, \ell = 2$ subshell would be labeled the 3d subshell.

When we assign electrons to states in a multielectron atom, we must be guided by the Pauli exclusion principle of Section 40-7; that is, no two electrons in an atom can have the same set of the quantum numbers n, ℓ, m_ℓ , and m_s . If this important principle did not hold, *all* the electrons in any atom could jump to the atom's lowest energy level, which would eliminate the chemistry of atoms and molecules, and thus also eliminate biochemistry and us. Let us examine the atoms

of a few elements to see how the Pauli exclusion principle operates in the building up of the periodic table.

Neon

The neon atom has 10 electrons. Only two of them fit into the lowest-energy subshell, the $1s$ subshell. These two electrons both have $n = 1$, $\ell = 0$, and $m_\ell = 0$, but one has $m_s = +\frac{1}{2}$ and the other has $m_s = -\frac{1}{2}$. The $1s$ subshell contains $2[2(0) + 1] = 2$ states. Because this subshell then contains all the electrons permitted by the Pauli principle, it is said to be **closed**.

Two of the remaining eight electrons fill the next lowest energy subshell, the $2s$ subshell. The last six electrons just fill the $2p$ subshell which, with $\ell = 1$, holds $2[2(1) + 1] = 6$ states.

In a closed subshell, all allowed z projections of the orbital angular momentum vector \vec{L} are present and, as you can verify from Fig. 40-5, these projections cancel for the subshell as a whole; for every positive projection there is a corresponding negative projection of the same magnitude. Similarly, the z projections of the spin angular momenta also cancel. Thus, a closed subshell has no angular momentum and no magnetic moment of any kind. Furthermore, its probability density is spherically symmetric. Then neon with its three closed subshells ($1s$, $2s$, and $2p$) has no “loosely dangling electrons” to encourage chemical interaction with other atoms. Neon, like the other **noble gases** that form the right-hand column of the periodic table, is almost chemically inert.

Sodium

Next after neon in the periodic table comes sodium, with 11 electrons. Ten of them form a closed neon-like core, which, as we have seen, has zero angular momentum. The remaining electron is largely outside this inert core, in the $3s$ subshell—the next lowest energy subshell. Because this **valence electron** of sodium is in a state with $\ell = 0$ (that is, an s state using the lettering system above), the sodium atom’s angular momentum and magnetic dipole moment must be due entirely to the spin of this single electron.

Sodium readily combines with other atoms that have a “vacancy” into which sodium’s loosely bound valence electron can fit. Sodium, like the other **alkali metals** that form the left-hand column of the periodic table, is chemically active.

Chlorine

The chlorine atom, which has 17 electrons, has a closed 10-electron, neon-like core, with 7 electrons left over. Two of them fill the $3s$ subshell, leaving five to be assigned to the $3p$ subshell, which is the subshell next lowest in energy. This subshell, which has $\ell = 1$, can hold $2[2(1) + 1] = 6$ electrons, and so there is a vacancy, or a “hole,” in this subshell.

Chlorine is receptive to interacting with other atoms that have a valence electron that might fill this hole. Sodium chloride (NaCl), for example, is a very stable compound. Chlorine, like the other **halogens** that form column VIIA of the periodic table, is chemically active.

Iron

The arrangement of the 26 electrons of the iron atom can be represented as follows:

$$\underline{1s^2 \quad 2s^2 \quad 2p^6 \quad 3s^2 \quad 3p^6 \quad 3d^6 \quad 4s^2}.$$

40-10 X RAYS AND THE ORDERING OF THE ELEMENTS

1127

The subshells are listed in numerical order and, following convention, a superscript gives the number of electrons in each subshell. From Table 40-1 we can see that an *s* subshell ($\ell = 0$) can hold 2 electrons, a *p* subshell ($\ell = 1$) can hold 6, and a *d* subshell ($\ell = 2$) can hold 10. Thus, iron's first 18 electrons form the five filled subshells that are marked off by the bracket, leaving 8 electrons to be accounted for. Six of the eight go into the *3d* subshell, and the remaining two go into the *4s* subshell.

The reason the last two electrons do not also go into the *3d* subshell (which can hold 10 electrons) is that the $3d^6\ 4s^2$ configuration results in a lower-energy state for the atom as a whole than would the $3d^8$ configuration. An iron atom with 8 electrons (rather than 6) in the *3d* subshell would quickly make a transition to the $3d^6\ 4s^2$ configuration, emitting electromagnetic radiation in the process. The lesson here is that except for the simplest elements, the states may not be filled in what we might think of as their "logical" sequence.

40-10 X Rays and the Ordering of the Elements

When a solid target, such as solid copper or tungsten, is bombarded with electrons whose kinetic energies are in the kiloelectron-volt range, electromagnetic radiation called **x rays** is emitted. Our concern here is what these rays—whose medical, dental, and industrial usefulness is so well known and widespread—can teach us about the atoms that absorb or emit them. Figure 40-13 shows the wavelength spectrum of the x rays produced when a beam of 35 keV electrons falls on a molybdenum target. We see a broad, continuous spectrum of radiation on which are superimposed two peaks of sharply defined wavelengths. The continuous spectrum and the peaks arise in different ways, which we next discuss separately.

The Continuous X-Ray Spectrum

Here we examine the continuous x-ray spectrum of Fig. 40-13, ignoring for the time being the two prominent peaks that rise from it. Consider an electron of initial kinetic energy K_0 that collides (interacts) with one of the target atoms, as in Fig. 40-14. The electron may lose an amount of energy ΔK , which will appear as the energy of an x-ray photon that is radiated away from the site of the collision. (Very little energy is transferred to the recoiling atom because of the relatively large mass of the atom; here we neglect that transfer.)

The scattered electron in Fig. 40-14, whose energy is now less than K_0 , may have a second collision with a target atom, generating a second photon, with a different photon energy. This electron-scattering process can continue until the electron is approximately stationary. All the photons generated by these collisions form part of the continuous x-ray spectrum.

A prominent feature of that spectrum in Fig. 40-13 is the sharply defined **cutoff wavelength** λ_{\min} , below which the continuous spectrum does not exist. This minimum wavelength corresponds to a collision in which an incident electron loses *all* its initial kinetic energy K_0 in a single head-on collision with a target atom. Essentially all this energy appears as the energy of a single photon, whose associated wavelength—the minimum possible x-ray wavelength—is found from

$$K_0 = hf = \frac{hc}{\lambda_{\min}},$$

or

$$\lambda_{\min} = \frac{hc}{K_0} \quad (\text{cutoff wavelength}). \quad (40-23)$$

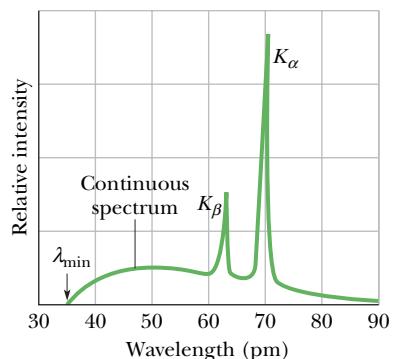


Fig. 40-13 The distribution by wavelength of the x rays produced when 35 keV electrons strike a molybdenum target. The sharp peaks and the continuous spectrum from which they rise are produced by different mechanisms.

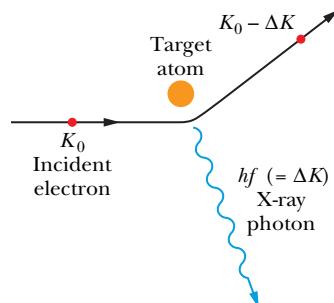


Fig. 40-14 An electron of kinetic energy K_0 passing near an atom in the target may generate an x-ray photon, the electron losing part of its energy in the process. The continuous x-ray spectrum arises in this way.

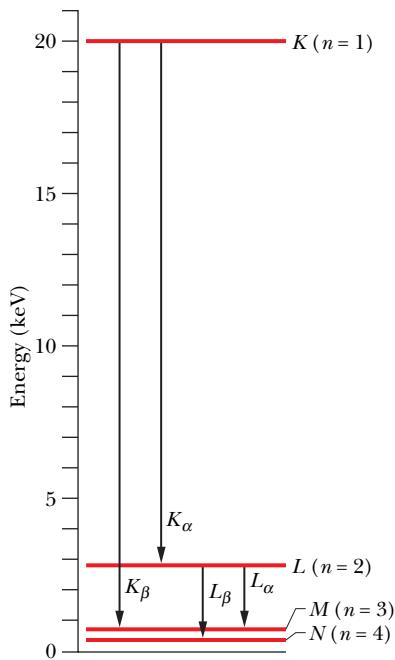


Fig. 40-15 A simplified energy-level diagram for a molybdenum atom, showing the transitions (of holes rather than electrons) that give rise to some of the characteristic x rays of that element. Each horizontal line represents the energy of the atom with a hole (a missing electron) in the shell indicated.

The cutoff wavelength is totally independent of the target material. If we were to switch from a molybdenum target to a copper target, for example, all features of the x-ray spectrum of Fig. 40-13 would change *except* the cutoff wavelength.

✓ CHECKPOINT 2

Does the cutoff wavelength λ_{\min} of the continuous x-ray spectrum increase, decrease, or remain the same if you (a) increase the kinetic energy of the electrons that strike the x-ray target, (b) allow the electrons to strike a thin foil rather than a thick block of the target material, (c) change the target to an element of higher atomic number?

The Characteristic X-Ray Spectrum

We now turn our attention to the two peaks of Fig. 40-13, labeled K_α and K_β . These (and other peaks that appear at wavelengths beyond the range displayed in Fig. 40-13) form the **characteristic x-ray spectrum** of the target material.

The peaks arise in a two-part process. (1) An energetic electron strikes an atom in the target and, while it is being scattered, the incident electron knocks out one of the atom's deep-lying (low n value) electrons. If the deep-lying electron is in the shell defined by $n = 1$ (called, for historical reasons, the K shell), there remains a vacancy, or *hole*, in this shell. (2) An electron in one of the shells with a higher energy jumps to the K shell, filling the hole in this shell. During this jump, the atom emits a characteristic x-ray photon. If the electron that fills the K -shell vacancy jumps from the shell with $n = 2$ (called the L shell), the emitted radiation is the K_α line of Fig. 40-13; if it jumps from the shell with $n = 3$ (called the M shell), it produces the K_β line, and so on. The hole left in either the L or M shell will be filled by an electron from still farther out in the atom.

In studying x rays, it is more convenient to keep track of where a hole is created deep in the atom's "electron cloud" than to record the changes in the quantum state of the electrons that jump to fill that hole. Figure 40-15 does exactly that; it is an energy-level diagram for molybdenum, the element to which Fig. 40-13 refers. The baseline ($E = 0$) represents the neutral atom in its ground state. The level marked K (at $E = 20$ keV) represents the energy of the molybdenum atom with a hole in its K shell, the level marked L (at $E = 2.7$ keV) represents the atom with a hole in its L shell, and so on.

The transitions marked K_α and K_β in Fig. 40-15 are the ones that produce the two x-ray peaks in Fig. 40-13. The K_α spectral line, for example, originates when an electron from the L shell fills a hole in the K shell. To state this transition in terms of what the arrows in Fig. 40-15 show, a hole originally in the K shell moves to the L shell.

Ordering the Elements

In 1913, British physicist H. G. J. Moseley generated characteristic x rays for as many elements as he could find—he found 38—by using them as targets for electron bombardment in an evacuated tube of his own design. By means of a trolley manipulated by strings, Moseley was able to move the individual targets into the path of an electron beam. He measured the wavelengths of the emitted x rays by the crystal diffraction method described in Section 36-10.

Moseley then sought (and found) regularities in these spectra as he moved from element to element in the periodic table. In particular, he noted that if, for a given spectral line such as K_α , he plotted for each element the square root of the frequency f against the position of the element in the periodic table, a straight line resulted. Figure 40-16 shows a portion of his extensive data. Moseley's conclusion was this:

We have here a proof that there is in the atom a fundamental quantity, which increases by regular steps as we pass from one element to the next. This quantity can only be the charge on the central nucleus.

As a result of Moseley's work, the characteristic x-ray spectrum became the universally accepted signature of an element, permitting the solution of a number of periodic table puzzles. Prior to that time (1913), the positions of elements in the table were assigned in order of atomic *mass*, although it was necessary to invert this order for several pairs of elements because of compelling chemical evidence; Moseley showed that it is the nuclear charge (that is, atomic number *Z*) that is the real basis for ordering the elements.

In 1913 the periodic table had several empty squares, and a surprising number of claims for new elements had been advanced. The x-ray spectrum provided a conclusive test of such claims. The lanthanide elements, often called the rare earth elements, had been sorted out only imperfectly because their similar chemical properties made sorting difficult. Once Moseley's work was reported, these elements were properly organized.

It is not hard to see why the characteristic x-ray spectrum shows such impressive regularities from element to element whereas the optical spectrum in the visible and near-visible region does not: The key to the identity of an element is the charge on its nucleus. Gold, for example, is what it is because its atoms have a nuclear charge of $+79e$ (that is, $Z = 79$). An atom with one more elementary charge on its nucleus is mercury; with one fewer, it is platinum. The *K* electrons, which play such a large role in the production of the x-ray spectrum, lie very close to the nucleus and are thus sensitive probes of its charge. The optical spectrum, on the other hand, involves transitions of the outermost electrons, which are heavily screened from the nucleus by the remaining electrons of the atom and thus are *not* sensitive probes of nuclear charge.

Accounting for the Moseley Plot

Moseley's experimental data, of which the Moseley plot of Fig. 40-16 is but a part, can be used directly to assign the elements to their proper places in the periodic table. This can be done even if no theoretical basis for Moseley's results can be established. However, there is such a basis.

According to Eqs. 39-32 and 39-33, the energy of the hydrogen atom is

$$E_n = -\frac{me^4}{8\varepsilon_0^2 h^2} \frac{1}{n^2} = -\frac{13.60 \text{ eV}}{n^2}, \quad \text{for } n = 1, 2, 3, \dots \quad (40-24)$$

Consider now one of the two innermost electrons in the *K* shell of a multi-electron atom. Because of the presence of the other *K*-shell electron, our electron "sees" an effective nuclear charge of approximately $(Z - 1)e$, where *e* is the

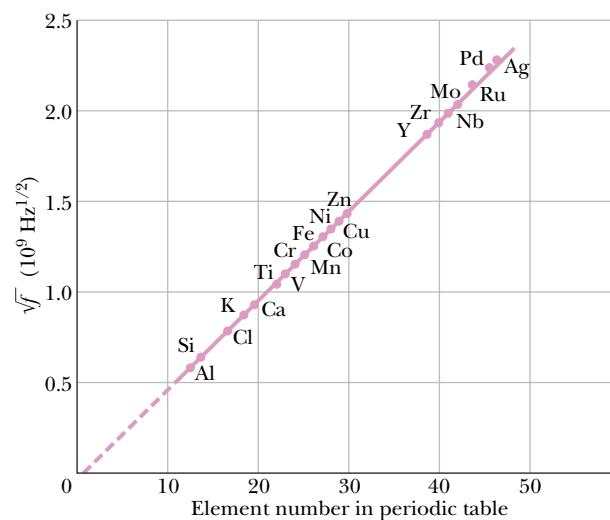


Fig. 40-16 A Moseley plot of the K_α line of the characteristic x-ray spectra of 21 elements. The frequency is calculated from the measured wavelength.

elementary charge and Z is the atomic number of the element. The factor e^4 in Eq. 40-24 is the product of e^2 —the square of hydrogen’s nuclear charge—and $(-e)^2$ —the square of an electron’s charge. For a multielectron atom, we can approximate the effective energy of the atom by replacing the factor e^4 in Eq. 40-24 with $(Z - 1)^2 e^2 \times (-e)^2$, or $e^4(Z - 1)^2$. That gives us

$$E_n = -\frac{(13.60 \text{ eV})(Z - 1)^2}{n^2}. \quad (40-25)$$

We saw that the K_α x-ray photon (of energy hf) arises when an electron makes a transition from the L shell (with $n = 2$ and energy E_2) to the K shell (with $n = 1$ and energy E_1). Thus, using Eq. 40-25, we may write the energy change as

$$\begin{aligned} \Delta E &= E_2 - E_1 \\ &= \frac{-(13.60 \text{ eV})(Z - 1)^2}{2^2} - \frac{-(13.60 \text{ eV})(Z - 1)^2}{1^2} \\ &= (10.2 \text{ eV})(Z - 1)^2. \end{aligned}$$

Then the frequency f of the K_α line is

$$\begin{aligned} f &= \frac{\Delta E}{h} = \frac{(10.2 \text{ eV})(Z - 1)^2}{(4.14 \times 10^{-15} \text{ eV} \cdot \text{s})} \\ &= (2.46 \times 10^{15} \text{ Hz})(Z - 1)^2. \end{aligned} \quad (40-26)$$

Taking the square root of both sides yields

$$\sqrt{f} = CZ - C, \quad (40-27)$$

in which C is a constant ($= 4.96 \times 10^7 \text{ Hz}^{1/2}$). Equation 40-27 is the equation of a straight line. It shows that if we plot the square root of the frequency of the K_α x-ray spectral line against the atomic number Z , we should obtain a straight line. As Fig. 40-16 shows, that is exactly what Moseley found.

Sample Problem

Characteristic spectrum in x-ray production

A cobalt target is bombarded with electrons, and the wavelengths of its characteristic x-ray spectrum are measured. There is also a second, fainter characteristic spectrum, which is due to an impurity in the cobalt. The wavelengths of the K_α lines are 178.9 pm (cobalt) and 143.5 pm (impurity), and the proton number for cobalt is $Z_{\text{Co}} = 27$. Determine the impurity using only these data.

KEY IDEA

The wavelengths of the K_α lines for both the cobalt (Co) and the impurity (X) fall on a K_α Moseley plot, and Eq. 40-27 is the equation for that plot.

Calculations: Substituting c/λ for f in Eq. 40-27, we obtain

$$\sqrt{\frac{c}{\lambda_{\text{Co}}}} = CZ_{\text{Co}} - C \quad \text{and} \quad \sqrt{\frac{c}{\lambda_X}} = CZ_X - C.$$

Dividing the second equation by the first neatly eliminates C , yielding

$$\sqrt{\frac{\lambda_{\text{Co}}}{\lambda_X}} = \frac{Z_X - 1}{Z_{\text{Co}} - 1}.$$

Substituting the given data yields

$$\sqrt{\frac{178.9 \text{ pm}}{143.5 \text{ pm}}} = \frac{Z_X - 1}{27 - 1}.$$

Solving for the unknown, we find that

$$Z_X = 30.0. \quad (\text{Answer})$$

Thus, the number of protons in the impurity nucleus is 30, and a glance at the periodic table identifies the impurity as zinc. Note that with a larger value of Z than cobalt, zinc has a smaller value of the K_α line. This means that the energy associated with that jump must be greater in zinc than cobalt.



Additional examples, video, and practice available at WileyPLUS

40-11 Lasers and Laser Light

In the early 1960s, quantum physics made one of its many contributions to technology: the **laser**. Laser light, like the light from an ordinary lightbulb, is emitted when atoms make a transition from one quantum state to a lower one. However, in a lightbulb the emissions are random, both in time and direction, and in a laser they are coordinated so that the emissions are at the same time and in the same direction. As a result, laser light has the following characteristics:

- Laser light is highly monochromatic.** Light from an ordinary incandescent lightbulb is spread over a continuous range of wavelengths and is certainly not monochromatic. The radiation from a fluorescent neon sign is monochromatic, true, to about 1 part in 10^6 , but the sharpness of definition of laser light can be many times greater, as much as 1 part in 10^{15} .
- Laser light is highly coherent.** Individual long waves (*wave trains*) for laser light can be several hundred kilometers long. When two separated beams that have traveled such distances over separate paths are recombined, they “remember” their common origin and are able to form a pattern of interference fringes. The corresponding *coherence length* for wave trains emitted by a lightbulb is typically less than a meter.
- Laser light is highly directional.** A laser beam spreads very little; it departs from strict parallelism only because of diffraction at the exit aperture of the laser. For example, a laser pulse used to measure the distance to the Moon generates a spot on the Moon’s surface with a diameter of only a few kilometers. Light from an ordinary bulb can be made into an approximately parallel beam by a lens, but the beam divergence is much greater than for laser light. Each point on a lightbulb’s filament forms its own separate beam, and the angular divergence of the overall composite beam is set by the size of the filament.
- Laser light can be sharply focused.** If two light beams transport the same amount of energy, the beam that can be focused to the smaller spot will have the greater intensity (power per unit area) at that spot. For laser light, the focused spot can be so small that an intensity of 10^{17} W/cm^2 is readily obtained. An oxyacetylene flame, by contrast, has an intensity of only about 10^3 W/cm^2 .

Lasers Have Many Uses

The smallest lasers, used for voice and data transmission over optical fibers, have as their active medium a semiconducting crystal about the size of a pinhead. Small as they are, such lasers can generate about 200 mW of power. The largest lasers, used for nuclear fusion research and for astronomical and military applications, fill a large building. The largest such laser can generate brief pulses of laser light with a power level, during the pulse, of about 10^{14} W . This is a few hundred times greater than the total electrical power generating capacity of the United States. To avoid a brief national power blackout during a pulse, the energy required for each pulse is stored up at a steady rate during the relatively long interpulse interval.

Among the many uses of lasers are reading bar codes, manufacturing and reading compact discs and DVDs, performing surgery of many kinds (both as a surgical aid as in Fig. 40-17 and as a cutting and cauterizing tool), surveying, cutting cloth in the garment industry (several hundred layers at a time), welding auto bodies, and generating holograms.



Fig. 40-17 A patient's head is scanned and mapped by (red) laser light in preparation for brain surgery. During the surgery, the laser-derived image of the head will be superimposed on the model of the brain shown on the monitor, to guide the surgical team into the region shown in green on the model. (Sam Ogden/Photo Researchers)

40-12 How Lasers Work

Because the word “laser” is an acronym for “light amplification by the stimulated emission of radiation,” you should not be surprised that stimulated emission is the key to laser operation. Einstein introduced this concept in 1917. Although the world had to wait until 1960 to see an operating laser, the groundwork for its development was put in place decades earlier.

Consider an isolated atom that can exist either in its state of lowest energy (its ground state), whose energy is E_0 , or in a state of higher energy (an excited state), whose energy is E_x . Here are three processes by which the atom can move from one of these states to the other:

- 1. Absorption.** Figure 40-18a shows the atom initially in its ground state. If the atom is placed in an electromagnetic field that is alternating at frequency f , the atom can absorb an amount of energy hf from that field and move to the higher-energy state. From the principle of conservation of energy we have

$$hf = E_x - E_0. \quad (40-28)$$

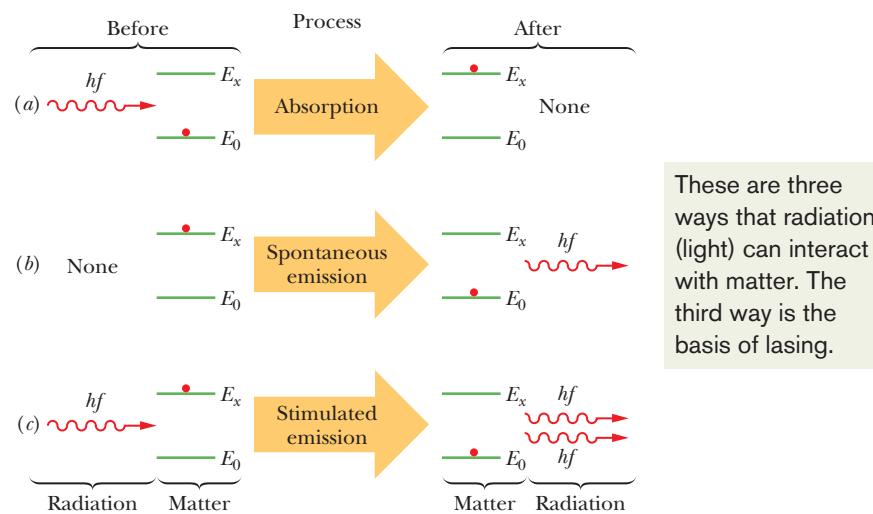
We call this process **absorption**.

- 2. Spontaneous emission.** In Fig. 40-18b the atom is in its excited state and no external radiation is present. After a time, the atom will move *of its own accord* to its ground state, emitting a photon of energy hf in the process. We call this process **spontaneous emission**—*spontaneous* because the event was not triggered by any outside influence. The light from the filament of an ordinary lightbulb is generated in this way.

Normally, the mean life of excited atoms before spontaneous emission occurs is about 10^{-8} s. However, for some excited states, this mean life is perhaps as much as 10^5 times longer. We call such long-lived states **metastable**; they play an important role in laser operation.

- 3. Stimulated emission.** In Fig. 40-18c the atom is again in its excited state, but this time radiation with a frequency given by Eq. 40-28 is present. A photon of energy hf can stimulate the atom to move to its ground state, during which process the atom emits an additional photon, whose energy is also hf . We call this process **stimulated emission**—*stimulated* because the event is triggered by the external photon. The emitted photon is in every way identical to the stimulating photon. Thus, the waves associated with the photons have the same energy, phase, polarization, and direction of travel.

Fig. 40-18 The interaction of radiation and matter in the processes of (a) absorption, (b) spontaneous emission, and (c) stimulated emission. An atom (matter) is represented by the red dot; the atom is in either a lower quantum state with energy E_0 or a higher quantum state with energy E_x . In (a) the atom absorbs a photon of energy hf from a passing light wave. In (b) it emits a light wave by emitting a photon of energy hf . In (c) a passing light wave with photon energy hf causes the atom to emit a photon of the same energy, increasing the energy of the light wave.



40-12 HOW LASERS WORK

1133

Figure 40-18c describes stimulated emission for a single atom. Suppose now that a sample contains a large number of atoms in thermal equilibrium at temperature T . Before any radiation is directed at the sample, a number N_0 of these atoms are in their ground state with energy E_0 and a number N_x are in a state of higher energy E_x . Ludwig Boltzmann showed that N_x is given in terms of N_0 by

$$N_x = N_0 e^{-(E_x - E_0)/kT}, \quad (40-29)$$

in which k is Boltzmann's constant. This equation seems reasonable. The quantity kT is the mean kinetic energy of an atom at temperature T . The higher the temperature, the more atoms—on average—will have been “bumped up” by thermal agitation (that is, by atom–atom collisions) to the higher energy state E_x . Also, because $E_x > E_0$, Eq. 40-29 requires that $N_x < N_0$; that is, there will always be fewer atoms in the excited state than in the ground state. This is what we expect if the level populations N_0 and N_x are determined only by the action of thermal agitation. Figure 40-19a illustrates this situation.

If we now flood the atoms of Fig. 40-19a with photons of energy $E_x - E_0$, photons will disappear via absorption by ground-state atoms and photons will be generated largely via stimulated emission of excited-state atoms. Einstein showed that the probabilities per atom for these two processes are identical. Thus, because there are more atoms in the ground state, the *net* effect will be the absorption of photons.

To produce laser light, we must have more photons emitted than absorbed; that is, we must have a situation in which stimulated emission dominates. Thus, we need more atoms in the excited state than in the ground state, as in Fig. 40-19b. However, because such a **population inversion** is not consistent with thermal equilibrium, we must think up clever ways to set up and maintain one.

The Helium–Neon Gas Laser

Figure 40-20 shows a common type of laser developed in 1961 by Ali Javan and his coworkers. The glass discharge tube is filled with a 20 : 80 mixture of helium and neon gases, neon being the medium in which laser action occurs.

Figure 40-21 shows simplified energy-level diagrams for the two types of atoms. An electric current passed through the helium–neon gas mixture serves—through

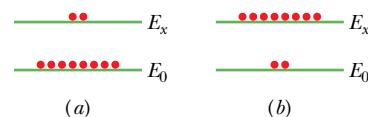


Fig. 40-19 (a) The equilibrium distribution of atoms between the ground state E_0 and excited state E_x accounted for by thermal agitation. (b) An inverted population, obtained by special methods. Such a population inversion is essential for laser action.

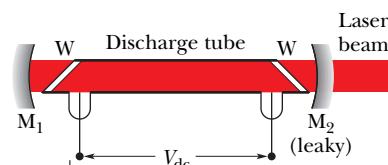


Fig. 40-20 The elements of a helium–neon gas laser. An applied potential V_{dc} sends electrons through a discharge tube containing a mixture of helium gas and neon gas. Electrons collide with helium atoms, which then collide with neon atoms, which emit light along the length of the tube. The light passes through transparent windows W and reflects back and forth through the tube from mirrors M_1 and M_2 to cause more neon atom emissions. Some of the light leaks through mirror M_2 to form the laser beam.

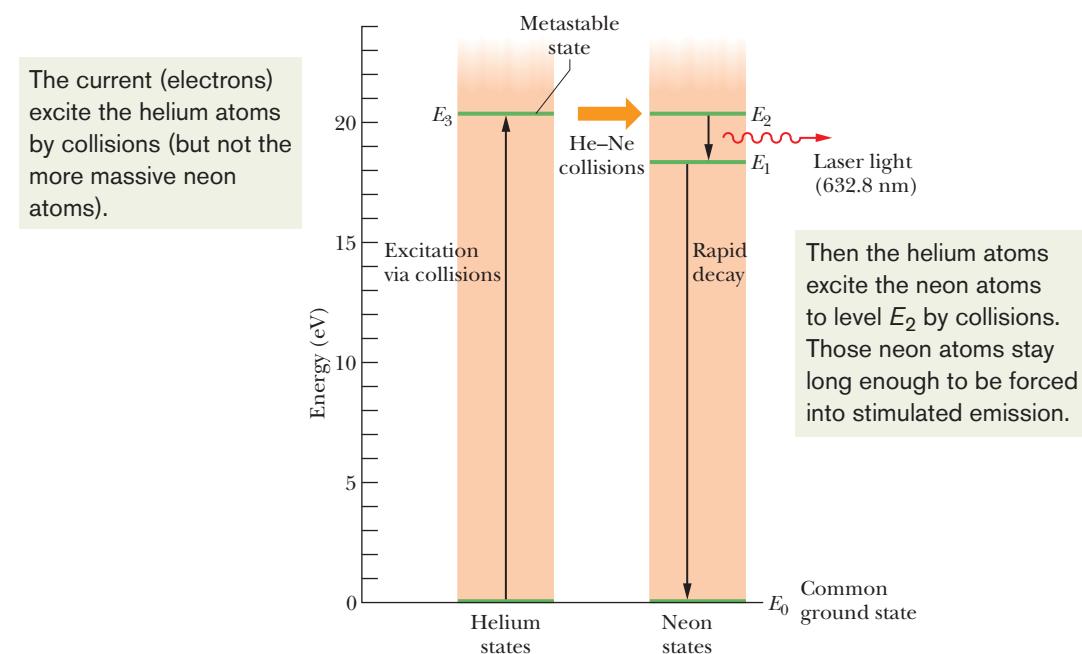


Fig. 40-21 Five essential energy levels for helium and neon atoms in a helium–neon gas laser. Laser action occurs between levels E_2 and E_1 of neon when more atoms are at the E_2 level than at the E_1 level.

collisions between helium atoms and electrons of the current—to raise many helium atoms to state E_3 , which is metastable with a mean life of at least 1 μs . (The neon atoms are too massive to be excited by collisions with the (low-mass) electrons.)

The energy of helium state E_3 (20.61 eV) is very close to the energy of neon state E_2 (20.66 eV). Thus, when a metastable (E_3) helium atom and a ground-state (E_0) neon atom collide, the excitation energy of the helium atom is often transferred to the neon atom, which then moves to state E_2 . In this manner, neon level E_2 (with a mean life of 170 ns) can become more heavily populated than neon level E_1 (which, with a mean life of only 10 ns, is almost empty).

This population inversion is relatively easy to set up because (1) initially there are essentially no neon atoms in state E_1 , (2) the long mean life of helium level E_3 means that there is always a good chance that collisions will excite neon atoms to their E_2 level, and (3) once those neon atoms undergo stimulated emission and fall to their E_1 level, they almost immediately fall down to their ground state (via intermediate levels not shown) and are then ready to be re-excited by collisions.

Suppose now that a single photon is spontaneously emitted as a neon atom transfers from state E_2 to state E_1 . Such a photon can trigger a stimulated emission event, which, in turn, can trigger other stimulated emission events. Through such a chain reaction, a coherent beam of laser light, moving parallel to the tube axis, can build up rapidly. This light, of wavelength 632.8 nm (red), moves through the discharge tube many times by successive reflections from mirrors M_1 and M_2 shown in Fig. 40-20, accumulating additional stimulated emission photons with each passage. M_1 is totally reflecting, but M_2 is slightly “leaky” so that a small fraction of the laser light escapes to form a useful external beam.



CHECKPOINT 3

The wavelength of light from laser *A* (a helium–neon gas laser) is 632.8 nm; that from laser *B* (a carbon dioxide gas laser) is 10.6 μm ; that from laser *C* (a gallium arsenide semiconductor laser) is 840 nm. Rank these lasers according to the energy interval between the two quantum states responsible for laser action, greatest first.

Sample Problem

Population inversion in a laser

In the helium–neon laser of Fig. 40-20, laser action occurs between two excited states of the neon atom. However, in many lasers, laser action (*lasing*) occurs between the ground state and an excited state, as suggested in Fig. 40-19*b*.

- (a) Consider such a laser that emits at wavelength $\lambda = 550$ nm. If a population inversion is not generated, what is the ratio of the population of atoms in state E_x to the population in the ground state E_0 , with the atoms at room temperature?

KEY IDEAS

- (1) The naturally occurring population ratio N_x/N_0 of the two states is due to thermal agitation of the gas atoms (Eq. 40-29):

$$N_x/N_0 = e^{-(E_x - E_0)/kT}. \quad (40-30)$$

To find N_x/N_0 with Eq. 40-30, we need to find the energy separation $E_x - E_0$ between the two states. (2) We can obtain $E_x - E_0$ from the given wavelength of 550 nm for the lasing between those two states.

Calculation: The lasing wavelength gives us

$$\begin{aligned} E_x - E_0 &= hf = \frac{hc}{\lambda} \\ &= \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s})}{(550 \times 10^{-9} \text{ m})(1.60 \times 10^{-19} \text{ J/eV})} \\ &= 2.26 \text{ eV}. \end{aligned}$$

To solve Eq. 40-30, we also need the mean energy of thermal agitation kT for an atom at room temperature (assumed to be 300 K), which is

$$kT = (8.62 \times 10^{-5} \text{ eV/K})(300 \text{ K}) = 0.0259 \text{ eV},$$

in which k is Boltzmann’s constant.

Substituting the last two results into Eq. 40-30 gives us the population ratio at room temperature:

$$\begin{aligned} N_x/N_0 &= e^{-(2.26 \text{ eV})/(0.0259 \text{ eV})} \\ &\approx 1.3 \times 10^{-38}. \end{aligned} \quad (\text{Answer})$$

This is an extremely small number. It is not unreasonable, however. Atoms with a mean thermal agitation energy of only 0.0259 eV will not often impart an energy of 2.26 eV to another atom in a collision.

- (b) For the conditions of (a), at what temperature would the ratio N_x/N_0 be 1/2?

Calculation: Now we want the temperature T such that thermal agitation has bumped enough neon atoms up to the higher-energy state to give $N_x/N_0 = 1/2$. Substituting that ratio into Eq. 40-30, taking the natural logarithm of both sides, and solving for T yield

$$\begin{aligned} T &= \frac{E_x - E_0}{k(\ln 2)} = \frac{2.26 \text{ eV}}{(8.62 \times 10^{-5} \text{ eV/K})(\ln 2)} \\ &= 38\,000 \text{ K}. \end{aligned} \quad (\text{Answer})$$

This is much hotter than the surface of the Sun. Thus, it is clear that if we are to invert the populations of these two levels, some specific mechanism for bringing this about is



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needed—that is, we must “pump” the atoms. No temperature, however high, will naturally generate a population inversion by thermal agitation.

REVIEW & SUMMARY

Some Properties of Atoms The energies of atoms are quantized; that is, the atoms have only certain specific values of energy associated with different quantum states. Atoms can make transitions between different quantum states by emitting or absorbing a photon; the frequency f associated with that light is given by

$$hf = E_{\text{high}} - E_{\text{low}}, \quad (40-1)$$

where E_{high} is the higher energy and E_{low} is the lower energy of the pair of quantum states involved in the transition. Atoms also have quantized angular momenta and magnetic dipole moments.

Angular Momenta and Magnetic Dipole Moments An electron trapped in an atom has an *orbital angular momentum* \vec{L} with a magnitude given by

$$L = \sqrt{\ell(\ell + 1)}\hbar, \quad (40-2)$$

where ℓ is the *orbital quantum number* (which can have the values given by Table 40-1) and where the constant “h-bar” is $\hbar = h/2\pi$. The projection L_z of \vec{L} on an arbitrary z axis is quantized and measurable and can have the values

$$L_z = m_\ell\hbar, \quad (40-7)$$

where m_ℓ is the *orbital magnetic quantum number* (which can have the values given by Table 40-1).

A magnetic dipole is associated with the angular momentum \vec{L} of an electron in an atom. This magnetic dipole has an **orbital magnetic dipole moment** $\vec{\mu}_{\text{orb}}$ that is directed opposite \vec{L} :

$$\vec{\mu}_{\text{orb}} = -\frac{e}{2m}\vec{L}, \quad (40-3)$$

where the minus sign indicates opposite directions. The projection $\mu_{\text{orb},z}$ of the orbital magnetic dipole moment on the z axis is quantized and measurable and can have the values

$$\mu_{\text{orb},z} = -m_\ell\mu_B, \quad (40-5)$$

where μ_B is the *Bohr magneton*:

$$\mu_B = \frac{e\hbar}{4\pi m} = 9.274 \times 10^{-24} \text{ J/T}. \quad (40-6)$$

An electron, whether trapped or free, has an intrinsic *spin angular momentum* (or just *spin*) \vec{S} with a magnitude given by

$$S = \sqrt{s(s + 1)}\hbar, \quad (40-9)$$

where s is the *spin quantum number* of the electron, which is always $\frac{1}{2}$. The projection S_z of \vec{S} on an arbitrary z axis is quantized

and measurable and can have the values

$$S_z = m_s\hbar, \quad (40-12)$$

where m_s is the *spin magnetic quantum number* of the electron, which can be $+\frac{1}{2}$ or $-\frac{1}{2}$.

An electron has an intrinsic magnetic dipole that is associated with its spin angular momentum \vec{S} , whether the electron is confined to an atom or free. This magnetic dipole has a **spin magnetic dipole moment** $\vec{\mu}_s$ that is directed opposite \vec{S} :

$$\vec{\mu}_s = -\frac{e}{m}\vec{S}. \quad (40-10)$$

The projection $\mu_{s,z}$ of the spin magnetic dipole moment $\vec{\mu}_s$ on an arbitrary z axis is quantized and measurable and can have the values

$$\mu_{s,z} = -2m_s\mu_B. \quad (40-13)$$

Spin and Magnetic Resonance A proton has an intrinsic spin angular momentum \vec{S} and an associated spin magnetic dipole moment $\vec{\mu}$ that is always in the *same* direction as \vec{S} . If a proton is located in magnetic field \vec{B} , the projection μ_z of $\vec{\mu}$ on a z axis (defined to be along the direction of \vec{B}) can have only two quantized orientations: parallel to \vec{B} or antiparallel to \vec{B} . The energy difference between these orientations is $2\mu_z B$. The energy required of a photon to *spin-flip* the proton between the two orientations is

$$hf = 2\mu_z B. \quad (40-22)$$

In general, \vec{B} is the vector sum of an external field \vec{B}_{ext} set up by the magnetic-resonance equipment and an internal field \vec{B}_{int} set up by the atoms and nuclei surrounding the proton. Detection of such spin-flips can lead to *nuclear magnetic resonance spectra* by which specific substances can be identified.

Pauli Exclusion Principle Electrons in atoms and other traps obey the **Pauli exclusion principle**, which requires that *no two electrons in a trap can have the same set of quantum numbers*.

Building the Periodic Table The elements are listed in the periodic table in order of increasing atomic number Z ; the nuclear charge is Ze , and Z is both the number of protons in the nucleus and the number of electrons in the neutral atom.

States with the same value of n form a **shell**, and those with the same values of both n and ℓ form a **subshell**. In *closed* shells and subshells, which are those that contain the maximum number of electrons, the angular momenta and the magnetic moments of the individual electrons sum to zero.

X Rays and the Numbering of the Elements A continuous spectrum of x rays is emitted when high-energy electrons lose some of their energy in a collision with atomic nuclei. The cutoff wavelength λ_{\min} is the wavelength emitted when such electrons lose all their initial energy in a single such encounter and is

$$\lambda_{\min} = \frac{hc}{K_0}, \quad (40-23)$$

in which K_0 is the initial kinetic energy of the electrons that strike the target.

The characteristic x-ray spectrum arises when high-energy electrons eject electrons from deep within the atom; when a resulting "hole" is filled by an electron from farther out in the atom, a photon of the characteristic x-ray spectrum is generated.

In 1913, British physicist H. G. J. Moseley measured the frequencies of the characteristic x rays from a number of elements. He noted that when the square root of the frequency is plotted against the position of the element in the periodic table, a straight

line results, as in the **Moseley plot** of Fig. 40-16. This allowed Moseley to conclude that the property that determines the position of an element in the periodic table is not its atomic mass (as was thought) but its atomic number Z —that is, the number of protons in its nucleus.

Lasers and Laser Light Laser light arises by **stimulated emission**. That is, radiation of a frequency given by

$$hf = E_x - E_0 \quad (40-28)$$

can cause an atom to undergo a transition from an upper energy level (of energy E_x) to a lower energy level, with a photon of frequency f being emitted. The stimulating photon and the emitted photon are identical in every respect (such as wavelength, phase, and direction of travel) and combine to form laser light.

For the emission process to predominate, there must normally be a **population inversion**; that is, there must be more atoms in the upper energy level than in the lower one.

Q U E S T I O N S

- 1 How many (a) subshells and (b) electron states are in the $n = 2$ shell? How many (c) subshells and (d) electron states are in the $n = 5$ shell?
- 2 An electron in an atom of gold is in a state with $n = 4$. Which of these values of ℓ are possible for it: $-3, 0, 2, 3, 4, 5$?
- 3 Label these statements as true or false: (a) One (and only one) of these subshells cannot exist: $2p, 4f, 3d, 1p$. (b) The number of values of m_ℓ that are allowed depends only on ℓ and not on n . (c) There are four subshells with $n = 4$. (d) The smallest value of n for a given value of ℓ is $\ell + 1$. (e) All states with $\ell = 0$ also have $m_\ell = 0$. (f) There are n subshells for each value of n .
- 4 An atom of uranium has closed $6p$ and $7s$ subshells. Which subshell has the greater number of electrons?
- 5 An atom of silver has closed $3d$ and $4d$ subshells. Which subshell has the greater number of electrons, or do they have the same number?
- 6 From which atom of each of the following pairs is it easier to remove an electron: (a) krypton or bromine, (b) rubidium or cerium, (c) helium or hydrogen?
- 7 An electron in a mercury atom is in the $3d$ subshell. Which of the following m_ℓ values are possible for it: $-3, -1, 0, 1, 2$?
- 8 Figure 40-22 shows three points at which a spin-up electron can be placed in a nonuniform magnetic field (there is a gradient along

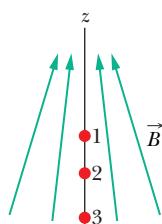


Fig. 40-22 Question 8.

the z axis). (a) Rank the three points according to the energy U of the electron's intrinsic magnetic dipole moment $\vec{\mu}_s$, most positive first. (b) What is the direction of the force on the electron due to the magnetic field if the spin-up electron is at point 2?

9 The K_α x-ray line for any element arises because of a transition between the K shell ($n = 1$) and the L shell ($n = 2$). Figure 40-13 shows this line (for a molybdenum target) occurring at a single wavelength. With higher resolution, however, the line splits into several wavelength components because the L shell does not have a unique energy. (a) How many components does the K_α line have? (b) Similarly, how many components does the K_β line have?

10 Consider the elements krypton and rubidium. (a) Which is more suitable for use in a Stern–Gerlach experiment of the kind described in connection with Fig. 40-8? (b) Which, if either, would not work at all?

11 On which quantum numbers does the energy of an electron depend in (a) a hydrogen atom and (b) a vanadium atom?

12 Which (if any) of the following are essential for laser action to occur between two energy levels of an atom? (a) There are more atoms in the upper level than in the lower. (b) The upper level is metastable. (c) The lower level is metastable. (d) The lower level is the ground state of the atom. (e) The lasing medium is a gas.

13 Figure 40-21 shows partial energy-level diagrams for the helium and neon atoms that are involved in the operation of a helium–neon laser. It is said that a helium atom in state E_3 can collide with a neon atom in its ground state and raise the neon atom to state E_2 . The energy of helium state E_3 (20.61 eV) is close to, but not exactly equal to, the energy of neon state E_2 (20.66 eV). How can the energy transfer take place if these energies are not *exactly* equal?

14 The x-ray spectrum of Fig. 40-13 is for 35.0 keV electrons striking a molybdenum ($Z = 42$) target. If you substitute a silver ($Z = 47$) target for the molybdenum target, will (a) λ_{\min} , (b) the wavelength for the K_α line, and (c) the wavelength for the K_β line increase, decrease, or remain unchanged?

PROBLEMS



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Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty

Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>**sec. 40-4 Angular Momenta and Magnetic Dipole Moments**

•1 An electron in a hydrogen atom is in a state with $\ell = 5$. What is the minimum possible value of the semiclassical angle between \vec{L} and L_z ?

•2 How many electron states are there in a shell defined by the quantum number $n = 5$?

•3 (a) What is the magnitude of the orbital angular momentum in a state with $\ell = 3$? (b) What is the magnitude of its largest projection on an imposed z axis?

•4 How many electron states are there in the following shells: (a) $n = 4$, (b) $n = 1$, (c) $n = 3$, (d) $n = 2$?

•5 (a) How many ℓ values are associated with $n = 3$? (b) How many m_ℓ values are associated with $\ell = 1$?

•6 How many electron states are in these subshells: (a) $n = 4$, $\ell = 3$; (b) $n = 3$, $\ell = 1$; (c) $n = 4$, $\ell = 1$; (d) $n = 2$, $\ell = 0$?

•7 An electron in a multielectron atom has $m_\ell = +4$. For this electron, what are (a) the value of ℓ , (b) the smallest possible value of n , and (c) the number of possible values of m_s ?

•8 In the subshell $\ell = 3$, (a) what is the greatest (most positive) m_ℓ value, (b) how many states are available with the greatest m_ℓ value, and (c) what is the total number of states available in the subshell?

•9 SSM WWW An electron is in a state with $\ell = 3$. (a) What multiple of \hbar gives the magnitude of \vec{L} ? (b) What multiple of μ_B gives the magnitude of $\vec{\mu}$? (c) What is the largest possible value of m_ℓ , (d) what multiple of \hbar gives the corresponding value of L_z , and (e) what multiple of μ_B gives the corresponding value of $\mu_{\text{orb},z}$? (f) What is the value of the semiclassical angle θ between the directions of L_z and \vec{L} ? What is the value of angle θ for (g) the second largest possible value of m_ℓ and (h) the smallest (that is, most negative) possible value of m_ℓ ?

•10 An electron is in a state with $n = 3$. What are (a) the number of possible values of ℓ , (b) the number of possible values of m_ℓ , (c) the number of possible values of m_s , (d) the number of states in the $n = 3$ shell, and (e) the number of subshells in the $n = 3$ shell?

•11 SSM If orbital angular momentum \vec{L} is measured along, say, a z axis to obtain a value for L_z , show that

$$(L_x^2 + L_y^2)^{1/2} = [\ell(\ell + 1) - m_\ell^2]^{1/2}\hbar$$

is the most that can be said about the other two components of the orbital angular momentum.

•12 GO A magnetic field is applied to a freely floating uniform iron sphere with radius $R = 2.00$ mm. The sphere initially had no net magnetic moment, but the field aligns 12% of the magnetic moments of the atoms (that is, 12% of the magnetic moments of the loosely bound electrons in the sphere, with one such electron per atom). The magnetic moment of those aligned electrons is the sphere's intrinsic magnetic moment $\vec{\mu}_s$. What is the sphere's resulting angular speed ω ?

sec. 40-5 The Stern-Gerlach Experiment

•13 SSM What is the acceleration of a silver atom as it passes through the deflecting magnet in the Stern-Gerlach experiment of Fig. 40-8 if the magnetic field gradient is 1.4 T/mm?

•14 Suppose that a hydrogen atom in its ground state moves 80 cm through and perpendicular to a vertical magnetic field that has a magnetic field gradient $dB/dz = 1.6 \times 10^2$ T/m. (a) What is the magnitude of force exerted by the field gradient on the atom due to the magnetic moment of the atom's electron, which we take to be 1 Bohr magneton? (b) What is the vertical displacement of the atom in the 80 cm of travel if its speed is 1.2×10^5 m/s?

•15 Calculate the (a) smaller and (b) larger value of the semiclassical angle between the electron spin angular momentum vector and the magnetic field in a Stern-Gerlach experiment. Bear in mind that the orbital angular momentum of the valence electron in the silver atom is zero.

•16 Assume that in the Stern-Gerlach experiment as described for neutral silver atoms, the magnetic field \vec{B} has a magnitude of 0.50 T. (a) What is the energy difference between the magnetic moment orientations of the silver atoms in the two subbeams? (b) What is the frequency of the radiation that would induce a transition between these two states? (c) What is the wavelength of this radiation, and (d) to what part of the electromagnetic spectrum does it belong?

sec. 40-6 Magnetic Resonance

•17 In an NMR experiment, the RF source oscillates at 34 MHz and magnetic resonance of the hydrogen atoms in the sample being investigated occurs when the external field \vec{B}_{ext} has magnitude 0.78 T. Assume that \vec{B}_{int} and \vec{B}_{ext} are in the same direction and take the proton magnetic moment component μ_z to be 1.41×10^{-26} J/T. What is the magnitude of \vec{B}_{int} ?

•18 A hydrogen atom in its ground state actually has two possible, closely spaced energy levels because the electron is in the magnetic field \vec{B} of the proton (the nucleus). Accordingly, an energy is associated with the orientation of the electron's magnetic moment $\vec{\mu}$ relative to \vec{B} , and the electron is said to be either spin up (higher energy) or spin down (lower energy) in that field. If the electron is excited to the higher-energy level, it can de-excite by spin-flipping and emitting a photon. The wavelength associated with that photon is 21 cm. (Such a process occurs extensively in the Milky Way galaxy, and reception of the 21 cm radiation by radio telescopes reveals where hydrogen gas lies between stars.) What is the effective magnitude of \vec{B} as experienced by the electron in the ground-state hydrogen atom?

•19 What is the wavelength associated with a photon that will induce a transition of an electron spin from parallel to antiparallel orientation in a magnetic field of magnitude 0.200 T? Assume that $\ell = 0$.

sec. 40-8 Multiple Electrons in Rectangular Traps

•20 A rectangular corral of widths $L_x = L$ and $L_y = 2L$ contains seven electrons. What multiple of $h^2/8mL^2$ gives the energy of the

1138 CHAPTER 40 ALL ABOUT ATOMS

ground state of this system? Assume that the electrons do not interact with one another, and do not neglect spin.

- 21 Seven electrons are trapped in a one-dimensional infinite potential well of width L . What multiple of $h^2/8mL^2$ gives the energy of the ground state of this system? Assume that the electrons do not interact with one another, and do not neglect spin.

•22 Figure 40-23 is an energy-level diagram for a fictitious infinite potential well of length L that contains one electron. The number of degenerate states of the levels are indicated: “non” means nondegenerate (which includes the ground state), “double” means 2 states, and “triple” means 3 states. We put a total of 11 electrons in the well. If the electrostatic forces between the electrons can be neglected, what multiple of $h^2/8mL^2$ gives the energy of the first excited state of the 11-electron system?

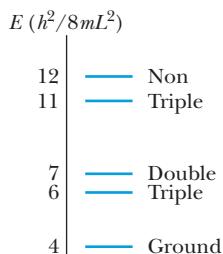


Fig. 40-23
Problem 22.

- 23 **SSM** A cubical box of widths $L_x = L_y = L_z = L$ contains eight electrons. What multiple of $h^2/8mL^2$ gives the energy of the ground state of this system? Assume that the electrons do not interact with one another, and do not neglect spin.

- 24 For Problem 20, what multiple of $h^2/8mL^2$ gives the energy of (a) the first excited state, (b) the second excited state, and (c) the third excited state of the system of seven electrons? (d) Construct an energy-level diagram for the lowest four energy levels.

- 25 For the situation of Problem 21, what multiple of $h^2/8mL^2$ gives the energy of (a) the first excited state, (b) the second excited state, and (c) the third excited state of the system of seven electrons? (d) Construct an energy-level diagram for the lowest four energy levels of the system.

- 26 For the situation of Problem 23, what multiple of $h^2/8mL^2$ gives the energy of (a) the first excited state, (b) the second excited state, and (c) the third excited state of the system of eight electrons? (d) Construct an energy-level diagram for the lowest four energy levels of the system.

sec. 40-9 Building the Periodic Table

- 27 **SSM WWW** Two of the three electrons in a lithium atom have quantum numbers (n, ℓ, m_ℓ, m_s) of $(1, 0, 0, +\frac{1}{2})$ and $(1, 0, 0, -\frac{1}{2})$. What quantum numbers are possible for the third electron if the atom is (a) in the ground state and (b) in the first excited state?

- 28 Show that the number of states with the same quantum number n is $2n^2$.

- 29 A recently named element is darmstadtium (Ds), which has 110 electrons. Assume that you can put the 110 electrons into the atomic shells one by one and can neglect any electron-electron interaction. With the atom in ground state, what is the spectroscopic notation for the quantum number ℓ for the last electron?

- 30 For a helium atom in its ground state, what are quantum numbers $(n, \ell, m_\ell, \text{ and } m_s)$ for the (a) spin-up electron and (b) spin-down electron?

- 31 Consider the elements selenium ($Z = 34$), bromine ($Z = 35$), and krypton ($Z = 36$). In their part of the periodic table, the subshells of the electronic states are filled in the sequence

$$1s \ 2s \ 2p \ 3s \ 3p \ 3d \ 4s \ 4p \ \dots$$

What are (a) the highest occupied subshell for selenium and (b) the number of electrons in it, (c) the highest occupied subshell for bromine and (d) the number of electrons in it, and (e) the highest occupied subshell for krypton and (f) the number of electrons in it?

- 32 Suppose two electrons in an atom have quantum numbers $n = 2$ and $\ell = 1$. (a) How many states are possible for those two electrons? (Keep in mind that the electrons are indistinguishable.) (b) If the Pauli exclusion principle did not apply to the electrons, how many states would be possible?

sec. 40-10 X Rays and the Ordering of the Elements

•33 Through what minimum potential difference must an electron in an x-ray tube be accelerated so that it can produce x rays with a wavelength of 0.100 nm?

- 34 The wavelength of the K_α line from iron is 193 pm. What is the energy difference between the two states of the iron atom that give rise to this transition?

- 35 **SSM WWW** In Fig. 40-13, the x rays shown are produced when 35.0 keV electrons strike a molybdenum ($Z = 42$) target. If the accelerating potential is maintained at this value but a silver ($Z = 47$) target is used instead, what values of (a) λ_{\min} , (b) the wavelength of the K_α line, and (c) the wavelength of the K_β line result? The K , L , and M atomic x-ray levels for silver (compare Fig. 40-15) are 25.51, 3.56, and 0.53 keV.

- 36 When electrons bombard a molybdenum target, they produce both continuous and characteristic x rays as shown in Fig. 40-13. In that figure the kinetic energy of the incident electrons is 35.0 keV. If the accelerating potential is increased to 50.0 keV, (a) what is the mean value of λ_{\min} , and (b) do the wavelengths of the K_α and K_β lines increase, decrease, or remain the same?

- 37 Show that a moving electron cannot spontaneously change into an x-ray photon in free space. A third body (atom or nucleus) must be present. Why is it needed? (Hint: Examine the conservation of energy and momentum.)

- 38 Here are the K_α wavelengths of a few elements:

Element	λ (pm)	Element	λ (pm)
Ti	275	Co	179
V	250	Ni	166
Cr	229	Cu	154
Mn	210	Zn	143
Fe	193	Ga	134

Make a Moseley plot (like that in Fig. 40-16) from these data and verify that its slope agrees with the value given for C in Section 40-10.

- 39 **SSM** Calculate the ratio of the wavelength of the K_α line for niobium (Nb) to that for gallium (Ga). Take needed data from the periodic table of Appendix G.

- 40 (a) From Eq. 40-26, what is the ratio of the photon energies due to K_α transitions in two atoms whose atomic numbers are Z and Z' ? (b) What is this ratio for uranium and aluminum? (c) For uranium and lithium?

- 41 The binding energies of K -shell and L -shell electrons in copper are 8.979 and 0.951 keV, respectively. If a K_α x ray from copper is incident on a sodium chloride crystal and gives a first-order

Bragg reflection at an angle of 74.1° measured relative to parallel planes of sodium atoms, what is the spacing between these parallel planes?

••42 From Fig. 40-13, calculate approximately the energy difference $E_L - E_M$ for molybdenum. Compare it with the value that may be obtained from Fig. 40-15.

••43 A tungsten ($Z = 74$) target is bombarded by electrons in an x-ray tube. The K , L , and M energy levels for tungsten (compare Fig. 40-15) have the energies 69.5, 11.3, and 2.30 keV, respectively. (a) What is the minimum value of the accelerating potential that will permit the production of the characteristic K_α and K_β lines of tungsten? (b) For this same accelerating potential, what is λ_{\min} ? What are the (c) K_α and (d) K_β wavelengths?

••44 A 20 keV electron is brought to rest by colliding twice with target nuclei as in Fig. 40-14. (Assume the nuclei remain stationary.) The wavelength associated with the photon emitted in the second collision is 130 pm greater than that associated with the photon emitted in the first collision. (a) What is the kinetic energy of the electron after the first collision? What are (b) the wavelength λ_1 and (c) the energy E_1 associated with the first photon? What are (d) λ_2 and (e) E_2 associated with the second photon?

••45 X rays are produced in an x-ray tube by electrons accelerated through an electric potential difference of 50.0 kV. Let K_0 be the kinetic energy of an electron at the end of the acceleration. The electron collides with a target nucleus (assume the nucleus remains stationary) and then has kinetic energy $K_1 = 0.500K_0$. (a) What wavelength is associated with the photon that is emitted? The electron collides with another target nucleus (assume it, too, remains stationary) and then has kinetic energy $K_2 = 0.500K_1$. (b) What wavelength is associated with the photon that is emitted?

••46 Determine the constant C in Eq. 40-27 to five significant figures by finding C in terms of the fundamental constants in Eq. 40-24 and then using data from Appendix B to evaluate those constants. Using this value of C in Eq. 40-27, determine the theoretical energy E_{theory} of the K_α photon for the low-mass elements listed in the following table. The table includes the value (eV) of the measured energy E_{exp} of the K_α photon for each listed element. The percentage deviation between E_{theory} and E_{exp} can be calculated as

$$\text{percentage deviation} = \frac{E_{\text{theory}} - E_{\text{exp}}}{E_{\text{exp}}} \times 100.$$

What is the percentage deviation for (a) Li, (b) Be, (c) B, (d) C, (e) N, (f) O, (g) F, (h) Ne, (i) Na, and (j) Mg?

Li	54.3	O	524.9
Be	108.5	F	676.8
B	183.3	Ne	848.6
C	277	Na	1041
N	392.4	Mg	1254

(There is actually more than one K_α ray because of the splitting of the L energy level, but that effect is negligible for the elements listed here.)

sec. 40-12 How Lasers Work

•47 The active volume of a laser constructed of the semiconductor GaAlAs is only $200 \mu\text{m}^3$ (smaller than a grain of sand),

and yet the laser can continuously deliver 5.0 mW of power at a wavelength of $0.80 \mu\text{m}$. At what rate does it generate photons?

•48 A high-powered laser beam ($\lambda = 600 \text{ nm}$) with a beam diameter of 12 cm is aimed at the Moon, $3.8 \times 10^5 \text{ km}$ distant. The beam spreads only because of diffraction. The angular location of the edge of the central diffraction disk (see Eq. 36-12) is given by

$$\sin \theta = \frac{1.22\lambda}{d},$$

where d is the diameter of the beam aperture. What is the diameter of the central diffraction disk on the Moon's surface?

•49 Assume that lasers are available whose wavelengths can be precisely “tuned” to anywhere in the visible range—that is, in the range $450 \text{ nm} < \lambda < 650 \text{ nm}$. If every television channel occupies a bandwidth of 10 MHz, how many channels can be accommodated within this wavelength range?

•50 A hypothetical atom has only two atomic energy levels, separated by 3.2 eV. Suppose that at a certain altitude in the atmosphere of a star there are $6.1 \times 10^{13}/\text{cm}^3$ of these atoms in the higher-energy state and $2.5 \times 10^{15}/\text{cm}^3$ in the lower-energy state. What is the temperature of the star's atmosphere at that altitude?

•51 **SSM** A hypothetical atom has energy levels uniformly separated by 1.2 eV. At a temperature of 2000 K, what is the ratio of the number of atoms in the 13th excited state to the number in the 11th excited state?

•52 **GO** A laser emits at 424 nm in a single pulse that lasts 0.500 μs . The power of the pulse is 2.80 MW. If we assume that the atoms contributing to the pulse underwent stimulated emission only once during the 0.500 μs , how many atoms contributed?

•53 A helium–neon laser emits laser light at a wavelength of 632.8 nm and a power of 2.3 mW. At what rate are photons emitted by this device?

•54 A certain gas laser can emit light at wavelength 550 nm, which involves population inversion between ground state and an excited state. At room temperature, how many moles of neon are needed to put 10 atoms in that excited state by thermal agitation?

•55 A pulsed laser emits light at a wavelength of 694.4 nm. The pulse duration is 12 ps, and the energy per pulse is 0.150 J. (a) What is the length of the pulse? (b) How many photons are emitted in each pulse?

•56 A population inversion for two energy levels is often described by assigning a negative Kelvin temperature to the system. What negative temperature would describe a system in which the population of the upper energy level exceeds that of the lower level by 10% and the energy difference between the two levels is 2.26 eV?

•57 A hypothetical atom has two energy levels, with a transition wavelength between them of 580 nm. In a particular sample at 300 K, 4.0×10^{20} such atoms are in the state of lower energy. (a) How many atoms are in the upper state, assuming conditions of thermal equilibrium? (b) Suppose, instead, that 3.0×10^{20} of these atoms are “pumped” into the upper state by an external process, with 1.0×10^{20} atoms remaining in the lower state. What is the maximum energy that could be released by the atoms in a single laser pulse if each atom jumps once between those two states (either via absorption or via stimulated emission)?

1140 CHAPTER 40 ALL ABOUT ATOMS

••58 The mirrors in the laser of Fig. 40-20, which are separated by 8.0 cm, form an optical cavity in which standing waves of laser light can be set up. Each standing wave has an integral number n of half wavelengths in the 8.0 cm length, where n is large and the waves differ slightly in wavelength. Near $\lambda = 533$ nm, how far apart in wavelength are the standing waves?

••59 Figure 40-24 shows the energy levels of two types of atoms. Atoms A are in one tube, and atoms B are in another tube. The energies (relative to a ground-state energy of zero) are indicated; the average lifetime of atoms in each level is also indicated. All the atoms are initially pumped to levels higher than the levels shown in the figure. The atoms then drop down through the levels, and many become “stuck” on certain levels, leading to population inversion and lasing. The light emitted by A illuminates B and can cause stimulated emission of B. What is the energy per photon of that stimulated emission of B?

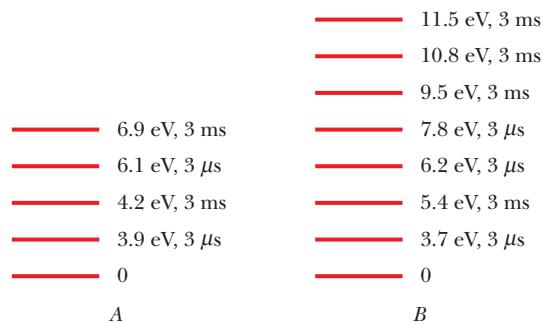


Fig. 40-24 Problem 59.

••60 The beam from an argon laser (of wavelength 515 nm) has a diameter d of 3.00 mm and a continuous energy output rate of 5.00 W. The beam is focused onto a diffuse surface by a lens whose focal length f is 3.50 cm. A diffraction pattern such as that of Fig. 36-10 is formed, the radius of the central disk being given by

$$R = \frac{1.22 f \lambda}{d}$$

(see Eq. 36-12 and Fig. 36-14). The central disk can be shown to contain 84% of the incident power. (a) What is the radius of the central disk? (b) What is the average intensity (power per unit area) in the incident beam? (c) What is the average intensity in the central disk?

••61 The active medium in a particular laser that generates laser light at a wavelength of 694 nm is 6.00 cm long and 1.00 cm in diameter. (a) Treat the medium as an optical resonance cavity analogous to a closed organ pipe. How many standing-wave nodes are there along the laser axis? (b) By what amount Δf would the beam frequency have to shift to increase this number by one? (c) Show that Δf is just the inverse of the travel time of laser light for one round trip back and forth along the laser axis. (d) What is the corresponding fractional frequency shift $\Delta f/f$? The appropriate index of refraction of the lasing medium (a ruby crystal) is 1.75.

••62 Ruby lasers at a wavelength of 694 nm. A certain ruby crystal has 4.00×10^{19} Cr ions (which are the atoms that lase). The lasing transition is between the first excited state and the ground state, and the output is a light pulse lasting 2.00 μ s. As the pulse begins, 60.0% of the Cr ions are in the first excited state and the rest are in the ground state. What is the average power emitted during the pulse? (Hint: Don't just ignore the ground-state ions.)

Additional Problems

63 Figure 40-25 is an energy-level diagram for a fictitious infinite potential well of length L that contains one electron. The number of degenerate states of the levels are indicated: “non” means non-degenerate (which includes the ground state) and “triple” means 3 states. If we put a total of 22 electrons in the well, what multiple of $h^2/8mL^2$ gives the energy of the ground state of the 22-electron system? Assume that the electrostatic forces between the electrons are negligible.

	$E (h^2/8mL^2)$
14	Triple
12	Non
11	Triple
9	Triple
6	Triple
3	Ground

Fig. 40-25
Problem 63.

64 *Martian CO₂ laser.* Where sunlight shines on the atmosphere of Mars, carbon dioxide molecules at an altitude of about 75 km undergo natural laser action. The energy levels involved in the action are shown in Fig. 40-26; population inversion occurs between energy levels E_2 and E_1 . (a) What wavelength of sunlight excites the molecules in the lasing action? (b) At what wavelength does lasing occur? (c) In what region of the electromagnetic spectrum do the excitation and lasing wavelengths lie?

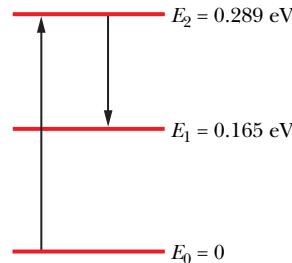


Fig. 40-26 Problem 64.

65 Excited sodium atoms emit two closely spaced spectrum lines called the *sodium doublet* (Fig. 40-27) with wavelengths 588.995 nm and 589.592 nm. (a) What is the difference in energy between the two upper energy levels ($n = 3, \ell = 1$)? (b) This energy difference occurs because the electron's spin magnetic moment can be oriented either parallel or antiparallel to the internal magnetic field associated with the electron's orbital motion. Use your result in (a) to find the magnitude of this internal magnetic field.

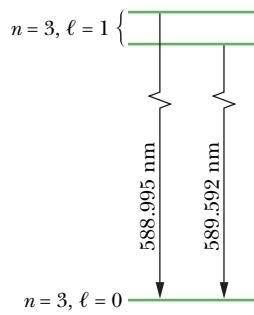


Fig. 40-27 Problem 65.

66 *Comet stimulated emission.* When a comet approaches the Sun, the increased warmth evaporates water from the ice on the surface of the comet nucleus, producing a thin atmosphere of water vapor around the nucleus. Sunlight can then dissociate H₂O molecules in the vapor to H atoms and OH molecules. The sunlight can also excite the OH molecules to higher energy levels.

When the comet is still relatively far from the Sun, the sunlight causes equal excitation to the E_2 and E_1 levels (Fig. 40-28a). Hence, there is no population inversion between the two levels. However, as the comet approaches the Sun, the excitation to the E_1 level decreases and population inversion occurs. The reason has to do with one of the many wavelengths—said to be *Fraunhofer lines*—that are missing in sunlight because, as the light travels outward through the Sun's atmosphere, those particular wavelengths are absorbed by the atmosphere.

As a comet approaches the Sun, the Doppler effect due to the comet's speed relative to the Sun shifts the Fraunhofer lines in wavelength, apparently overlapping one of them with the wavelength required for excitation to the E_1 level in OH molecules. Population inversion then occurs in those molecules, and they radiate stimulated emission (Fig. 40-28b). For example, as comet Kohoutek approached the Sun in December 1973 and January 1974, it radiated stimulated emission at about 1666 MHz during mid-January. (a) What was the energy difference $E_2 - E_1$ for that emission? (b) In what region of the electromagnetic spectrum was the emission?

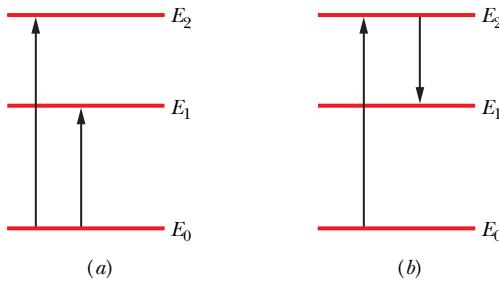


Fig. 40-28 Problem 66.

67 Show that the cutoff wavelength (in picometers) in the continuous x-ray spectrum from any target is given by $\lambda_{\min} = 1240/V$, where V is the potential difference (in kilovolts) through which the electrons are accelerated before they strike the target.

68 By measuring the go-and-return time for a laser pulse to travel from an Earth-bound observatory to a reflector on the Moon, it is possible to measure the separation between these bodies. (a) What is the predicted value of this time? (b) The separation can be measured to a precision of about 15 cm. To what uncertainty in travel time does this correspond? (c) If the laser beam forms a spot on the Moon 3 km in diameter, what is the angular divergence of the beam?

69 SSM Can an incoming intercontinental ballistic missile be destroyed by an intense laser beam? A beam of intensity 10^8 W/m^2 would probably burn into and destroy a nonspinning missile in 1 s. (a) If the laser had 5.0 MW power, 3.0 μm wavelength, and a 4.0 m beam diameter (a very powerful laser indeed), would it destroy a missile at a distance of 3000 km? (b) If the wavelength could be

changed, what maximum value would work? Use the equation for the central diffraction maximum as given by Eq. 36-12 ($\sin \theta = 1.22\lambda/d$).

70 A molybdenum ($Z = 42$) target is bombarded with 35.0 keV electrons and the x-ray spectrum of Fig. 40-13 results. The K_β and K_α wavelengths are 63.0 and 71.0 pm, respectively. What photon energy corresponds to the (a) K_β and (b) K_α radiation? The two radiations are to be filtered through one of the substances in the following table such that the substance absorbs the K_β line more strongly than the K_α line. A substance will absorb radiation x_1 more strongly than it absorbs radiation x_2 if a photon of x_1 has enough energy to eject a K electron from an atom of the substance but a photon of x_2 does not. The table gives the ionization energy of the K electron in molybdenum and four other substances. Which substance in the table will serve (c) best and (d) second best as the filter?

	Zr	Nb	Mo	Tc	Ru
Z	40	40	42	43	44
$E_K (\text{keV})$	18.00	18.99	20.00	21.04	22.12

71 An electron in a multielectron atom is known to have the quantum number $\ell = 3$. What are its possible n , m_ℓ , and m_s quantum numbers?

72 Show that if the 63 electrons in an atom of europium were assigned to shells according to the “logical” sequence of quantum numbers, this element would be chemically similar to sodium.

73 SSM Lasers can be used to generate pulses of light whose durations are as short as 10 fs. (a) How many wavelengths of light ($\lambda = 500 \text{ nm}$) are contained in such a pulse? (b) In

$$\frac{10 \text{ fs}}{1 \text{ s}} = \frac{1 \text{ s}}{X},$$

what is the missing quantity X (in years)?

74 Show that $\hbar = 1.06 \times 10^{-34} \text{ J}\cdot\text{s} = 6.59 \times 10^{-16} \text{ eV}\cdot\text{s}$.

75 Suppose that the electron had no spin and that the Pauli exclusion principle still held. Which, if any, of the present noble gases would remain in that category?

76 (A correspondence principle problem.) Estimate (a) the quantum number ℓ for the orbital motion of Earth around the Sun and (b) the number of allowed orientations of the plane of Earth's orbit. (c) Find θ_{\min} , the half-angle of the smallest cone that can be swept out by a perpendicular to Earth's orbit as Earth revolves around the Sun.

77 Knowing that the minimum x-ray wavelength produced by 40.0 keV electrons striking a target is 31.1 pm, determine the Planck constant \hbar .

41

CONDUCTION OF ELECTRICITY IN SOLIDS

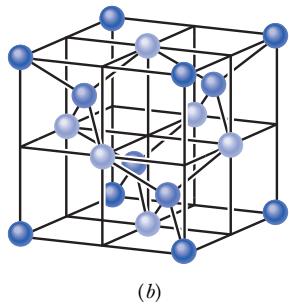
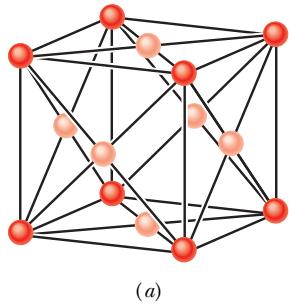


Fig. 41-1 (a) The unit cell for copper is a cube. There is one copper atom (darker) at each corner of the cube and one copper atom (lighter) at the center of each face of the cube. The arrangement is called *face-centered cubic*. (b) The unit cell for either silicon or the carbon atoms in diamond is also a cube, the atoms being arranged in what is called a *diamond lattice*. There is one atom (darkest) at each corner of the cube and one atom (lightest) at the center of each cube face; in addition, four atoms (medium color) lie within the cube. Every atom is bonded to its four nearest neighbors by a two-electron covalent bond (only the four atoms within the cube show all four nearest neighbors).

41-1 WHAT IS PHYSICS?

In this chapter we focus on a goal of physics that has become enormously important in the last half-century. That goal is to answer the question: What are the mechanisms by which a material conducts, or does not conduct, electricity? The answers to that question are complex and still not well understood, largely because they involve the application of quantum physics—not to individual particles and atoms as we have seen in the last several chapters but to a tremendous number of particles and atoms grouped together and interacting. In spite of the difficulty, engineers and scientists have made great strides in the quantum physics of materials science, which is why we now have computers, calculators, cell phones, and many other types of *solid-state* electronic devices.

Our starting point in this chapter is to characterize the solids that can conduct electricity and those that cannot.

41-2 The Electrical Properties of Solids

We shall examine only **crystalline solids**—that is, solids whose atoms are arranged in a repetitive three-dimensional structure called a **lattice**. We shall not consider such solids as wood, plastic, glass, and rubber, whose atoms are not arranged in such repetitive patterns. Figure 41-1 shows the basic repetitive units (the **unit cells**) of the lattice structures of copper, our prototype of a metal, and silicon and diamond (carbon), our prototypes of a semiconductor and an insulator, respectively.

We can classify solids electrically according to three basic properties:

1. Their **resistivity** ρ at room temperature, with the SI unit ohm-meter ($\Omega \cdot m$); resistivity is defined in Section 26-4.
2. Their **temperature coefficient of resistivity** α , defined as $\alpha = (1/\rho)(d\rho/dT)$ in Eq. 26-17 and having the SI unit inverse kelvin (K^{-1}). We can evaluate α for any solid by measuring ρ over a range of temperatures.
3. Their **number density of charge carriers** n . This quantity, the number of charge carriers per unit volume, can be found from measurements of the Hall effect, as discussed in Section 28-4, and from other measurements. It has the SI unit inverse cubic meter (m^{-3}).

From measurements of room-temperature resistivity alone, we discover that there are some materials—we call them **insulators**—that for all practical purposes do not conduct electricity at all. These are materials with very high resistivity. Diamond, an excellent example, has a resistivity greater than that of copper by the enormous factor of about 10^{24} .

41-3 ENERGY LEVELS IN A CRYSTALLINE SOLID

1143

We can then use measurements of ρ , α , and n to divide most noninsulators, at least at low temperatures, into two categories: **metals** and **semiconductors**.

Semiconductors have a considerably greater resistivity ρ than metals.

Semiconductors have a temperature coefficient of resistivity α that is both high and negative. That is, the resistivity of a semiconductor *decreases* with temperature, whereas that of a metal *increases*.

Semiconductors have a considerably lower number density of charge carriers n than metals.

Table 41-1 shows values of these quantities for copper, our prototype metal, and silicon, our prototype semiconductor.

Now let's consider our central question: *What features make diamond an insulator, copper a metal, and silicon a semiconductor?*

41-3 Energy Levels in a Crystalline Solid

The distance between adjacent copper atoms in solid copper is 260 pm. Figure 41-2a shows two isolated copper atoms separated by a distance r that is much greater than that. As Fig. 41-2b shows, each of these isolated neutral atoms stacks up its 29 electrons in an array of discrete subshells as follows:

$$1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 3d^{10} \ 4s^1.$$

Here we use the shorthand notation of Section 40-9 to identify the subshells. Recall, for example, that the subshell with principal quantum number $n = 3$ and orbital quantum number $\ell = 1$ is called the $3p$ subshell; it can hold up to $2(2\ell + 1) = 6$ electrons; the number it actually contains is indicated by a numerical superscript. We see above that the first six subshells in copper are filled, but the (outermost) $4s$ subshell, which can hold two electrons, holds only one.

If we bring the atoms of Fig. 41-2a closer together, their wave functions will start to overlap, beginning with those of the outermost electrons. When the wave functions of the two atoms overlap, we speak not of two independent atoms but of a single two-atom system; here the system contains $2 \times 29 = 58$ electrons. The Pauli exclusion principle also applies to this larger system and requires that each of these 58 electrons occupy a different quantum state. In fact, 58 quantum states are available because each energy level of the isolated atom splits into *two* levels for the two-atom system.

If we bring up more atoms, we gradually assemble a lattice of solid copper. If our lattice contains, say, N atoms, then each level of an isolated copper atom must split into N levels in the solid. Thus, the individual energy levels of the solid form energy **bands**, adjacent bands being separated by an energy **gap**, with the gap representing a range of energies that no electron can possess. A typical band ranges over only a few electron-volts. Since N may be of the order of 10^{24} , we see that the individual levels within a band are very close together indeed, and there are a vast number of levels.

Table 41-1

Some Electrical Properties of Two Materials^a

Property	Unit	Material	
		Copper	Silicon
Type of conductor		Metal	Semiconductor
Resistivity, ρ	$\Omega \cdot m$	2×10^{-8}	3×10^3
Temperature coefficient of resistivity, α	K^{-1}	$+4 \times 10^{-3}$	-70×10^{-3}
Number density of charge carriers, n	m^{-3}	9×10^{28}	1×10^{16}

^aAll values are for room temperature.

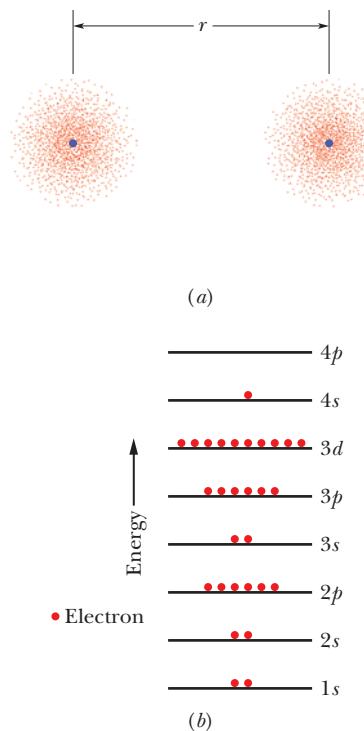


Fig. 41-2 (a) Two copper atoms separated by a large distance; their electron distributions are represented by dot plots. (b) Each copper atom has 29 electrons distributed among a set of subshells. In the neutral atom in its ground state, all subshells up through the $3d$ level are filled, the $4s$ subshell contains one electron (it can hold two), and higher subshells are empty. For simplicity, the subshells are shown as being evenly spaced in energy.

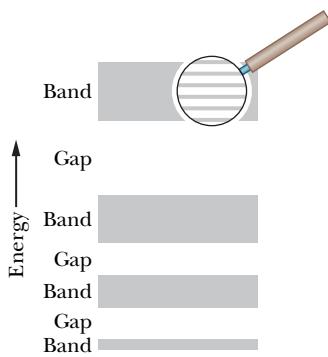


Fig. 41-3 The band–gap pattern of energy levels for an idealized crystalline solid. As the magnified view suggests, each band consists of a very large number of very closely spaced energy levels. (In many solids, adjacent bands may overlap; for clarity, we have not shown this condition.)

Figure 41-3 suggests the band–gap structure of the energy levels in a generalized crystalline solid. Note that bands of lower energy are narrower than those of higher energy. This occurs because electrons that occupy the lower energy bands spend most of their time deep within the atom's electron cloud. The wave functions of these core electrons do not overlap as much as the wave functions of the outer electrons do. Hence the splitting of the lower energy levels (core electrons) is less than that of the higher energy levels (outer electrons).

41-4 Insulators

A solid is said to be an electrical insulator if no current exists within it when we apply a potential difference across it. For a current to exist, the kinetic energy of the average electron must increase. In other words, some electrons in the solid must move to a higher energy level. However, as Fig. 41-4 shows, in an insulator the highest band containing any electrons is fully occupied. Because the Pauli exclusion principle keeps electrons from moving to occupied levels, no electrons in the solid are allowed to move. Thus, the electrons in the filled band of an insulator have no place to go; they are in gridlock, like a child on a ladder filled with children.

There are plenty of unoccupied levels (or *vacant levels*) in the band above the filled band in Fig. 41-4. However, if an electron is to occupy one of those levels, it must acquire enough energy to jump across the substantial energy gap E_g that separates the two bands. In diamond, this gap is so wide (the energy needed to cross it is 5.5 eV, about 140 times the average thermal energy of a free particle at room temperature) that essentially no electron can jump across it. Diamond is thus an electrical insulator, and a very good one.

In an insulator,
electrons need
a big energy
jump.

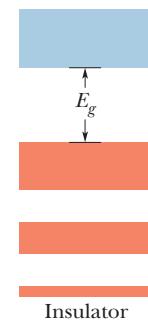


Fig. 41-4 The band–gap pattern for an insulator; filled levels are shown in red and empty levels in blue.

Sample Problem

Probability of electron excitation in an insulator

Approximately what is the probability that, at room temperature (300 K), an electron at the top of the highest filled band in diamond (an insulator) will jump the energy gap E_g in Fig. 41-4? For diamond, E_g is 5.5 eV.

KEY IDEA

In Chapter 40 we used Eq. 40-29,

$$\frac{N_x}{N_0} = e^{-(E_x - E_0)/kT}, \quad (41-1)$$

to relate the population N_x of atoms at energy level E_x to the population N_0 at energy level E_0 , where the atoms are part of a system at temperature T (measured in kelvins); k is the Boltzmann constant (8.62×10^{-5} eV/K). In this chapter we can use Eq. 41-1 to *approximate* the probability P that an electron in an insulator will jump the energy gap E_g in Fig. 41-4.

Calculations: We first set the energy difference $E_x - E_0$ to E_g . Then the probability P of the jump is approximately equal to the ratio N_x/N_0 of the number of electrons just above the energy gap to the number of electrons just below the gap.

For diamond, the exponent in Eq. 41-1 is

$$-\frac{E_g}{kT} = -\frac{5.5 \text{ eV}}{(8.62 \times 10^{-5} \text{ eV/K})(300 \text{ K})} = -213.$$

The required probability is then

$$P = \frac{N_x}{N_0} = e^{-(E_g/kT)} = e^{-213} \approx 3 \times 10^{-93}. \quad (\text{Answer})$$

This result tells us that approximately 3 electrons out of 10^{93} electrons would jump across the energy gap. Because any diamond stone has fewer than 10^{23} electrons, we see that the probability of the jump is vanishingly small. No wonder diamond is such a good insulator.



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41-5 Metals

The feature that defines a metal is that, as Fig. 41-5 shows, the highest occupied energy level falls somewhere near the middle of an energy band. If we apply a potential difference across a metal, a current can exist because there are plenty of vacant levels at nearby higher energies into which electrons (the charge carriers in a metal) can jump. Thus, a metal can conduct electricity because electrons in its highest occupied band can easily move into higher energy levels.

In Section 26-6 we discussed the **free-electron model** of a metal, in which the **conduction electrons** are free to move throughout the volume of the sample like the molecules of a gas in a closed container. We used this model to derive an expression for the resistivity of a metal. Here we use the model to explain the behavior of the conduction electrons in the partially filled band of Fig. 41-5. However, we now assume the energies of these electrons to be quantized and the Pauli exclusion principle to hold.

Assuming that the electric potential energy U of a conduction electron is uniform throughout the lattice, let's set $U = 0$ so that the mechanical energy E is entirely kinetic. Then the level at the bottom of the partially filled band of Fig. 41-5 corresponds to $E = 0$. The highest occupied level in this band at absolute zero ($T = 0$ K) is called the **Fermi level**, and the energy corresponding to it is called the **Fermi energy** E_F ; for copper, $E_F = 7.0$ eV.

The electron speed corresponding to the Fermi energy is called the **Fermi speed** v_F . For copper the Fermi speed is 1.6×10^6 m/s. Thus, all motion does *not* cease at absolute zero; at that temperature—and solely because of the Pauli exclusion principle—the conduction electrons are stacked up in the partially filled band of Fig. 41-5 with energies that range from zero to the Fermi energy.

How Many Conduction Electrons Are There?

If we could bring individual atoms together to form a sample of a metal, we would find that the conduction electrons in the metal are the *valence electrons* of the atoms (the electrons in the outermost occupied shells of the atoms). A *monovalent* atom contributes one such electron to the conduction electrons in a metal; a *bivalent* atom contributes two such electrons. Thus, the total number of conduction electrons is

$$\left(\begin{array}{l} \text{number of conduction} \\ \text{electrons in sample} \end{array} \right) = \left(\begin{array}{l} \text{number of atoms} \\ \text{in sample} \end{array} \right) \left(\begin{array}{l} \text{number of valence} \\ \text{electrons per atom} \end{array} \right). \quad (41-2)$$

(In this chapter, we shall write several equations largely in words because the symbols we have previously used for the quantities in them now represent other quantities.) The *number density* n of conduction electrons in a sample is the number of conduction electrons per unit volume:

$$n = \frac{\text{number of conduction electrons in sample}}{\text{sample volume } V}. \quad (41-3)$$

We can relate the number of atoms in a sample to various other properties of the sample and to the material making up the sample with the following:

$$\begin{aligned} \left(\begin{array}{l} \text{number of atoms} \\ \text{in sample} \end{array} \right) &= \frac{\text{sample mass } M_{\text{sam}}}{\text{atomic mass}} = \frac{\text{sample mass } M_{\text{sam}}}{(\text{molar mass } M)/N_A} \\ &= \frac{(\text{material's density})(\text{sample volume } V)}{(\text{molar mass } M)/N_A}, \end{aligned} \quad (41-4)$$

where the molar mass M is the mass of one mole of the material in the sample and N_A is Avogadro's number (6.02×10^{23} mol $^{-1}$).

In a conductor, electrons need only a small energy jump.

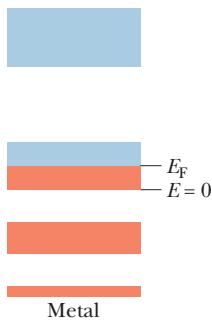


Fig. 41-5 The band-gap pattern for a metal. The highest filled level, called the Fermi level, lies near the middle of a band. Since vacant levels are available within that band, electrons in the band can easily change levels, and conduction can take place.

Sample Problem**Number of conduction electrons in a metal**

How many conduction electrons are in a cube of magnesium of volume $2.00 \times 10^{-6} \text{ m}^3$? Magnesium atoms are bivalent.

KEY IDEAS

- Because magnesium atoms are bivalent, each magnesium atom contributes two conduction electrons.
- The cube's number of conduction electrons is related to its number of magnesium atoms by Eq. 41-2.
- We can find the number of atoms with Eq. 41-4 and known data about the cube's volume and magnesium's properties.

Calculations: We can write Eq. 41-4 as

$$\left(\begin{array}{l} \text{number} \\ \text{of atoms} \\ \text{in sample} \end{array} \right) = \frac{(\text{density})(\text{sample volume } V)N_A}{\text{molar mass } M}.$$

Magnesium has density 1.738 g/cm^3 ($= 1.738 \times 10^3 \text{ kg/m}^3$)

and molar mass 24.312 g/mol ($= 24.312 \times 10^{-3} \text{ kg/mol}$) (see Appendix F). The numerator gives us

$$(1.738 \times 10^3 \text{ kg/m}^3)(2.00 \times 10^{-6} \text{ m}^3) \times (6.02 \times 10^{23} \text{ atoms/mol}) = 2.0926 \times 10^{21} \text{ kg/mol.}$$

$$\text{Thus, } \left(\begin{array}{l} \text{number of atoms} \\ \text{in sample} \end{array} \right) = \frac{2.0926 \times 10^{21} \text{ kg/mol}}{24.312 \times 10^{-3} \text{ kg/mol}} = 8.61 \times 10^{22}.$$

Using this result and the fact that magnesium atoms are bivalent, we find that Eq. 41-2 yields

$$\left(\begin{array}{l} \text{number of} \\ \text{conduction electrons} \\ \text{in sample} \end{array} \right) = (8.61 \times 10^{22} \text{ atoms}) \left(2 \frac{\text{electrons}}{\text{atom}} \right) = 1.72 \times 10^{23} \text{ electrons.} \quad (\text{Answer})$$



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Conductivity Above Absolute Zero

Our practical interest in the conduction of electricity in metals is at temperatures above absolute zero. What happens to the electron distribution of Fig. 41-5 at such higher temperatures? As we shall see, surprisingly little. Of the electrons in the partially filled band of Fig. 41-5, only those that are close to the Fermi energy find unoccupied levels above them, and only those electrons are free to be boosted to these higher levels by thermal agitation. Even at $T = 1000 \text{ K}$ (the copper would glow brightly in a dark room), the electron distribution among the available levels does not differ much from the distribution at $T = 0 \text{ K}$.

Let us see why. The quantity kT , where k is the Boltzmann constant, is a convenient measure of the energy that may be given to a conduction electron by the random thermal motions of the lattice. At $T = 1000 \text{ K}$, we have $kT = 0.086 \text{ eV}$. No electron can hope to have its energy changed by more than a few times this relatively small amount by thermal agitation alone; so at best only those few conduction electrons whose energies are close to the Fermi energy are likely to jump to higher energy levels due to thermal agitation. Poetically stated, thermal agitation normally causes only ripples on the surface of the Fermi sea of electrons; the vast depths of that sea lie undisturbed.

How Many Quantum States Are There?

The ability of a metal to conduct electricity depends on how many quantum states are available to its electrons and what the energies of those states are. Thus, a question arises: What are the energies of the individual states in the partially filled band of Fig. 41-5? This question is too difficult to answer because we cannot possibly list the energies of so many states individually. We ask instead: How many states in a unit volume of a sample have energies in the energy range E to $E + dE$? We write this number as $N(E) dE$, where $N(E)$ is called the **density of states** at energy E . The conventional unit for $N(E) dE$ is states per cubic meter (states/ m^3 , or simply m^{-3}), and the conventional unit for $N(E)$ is states per cubic meter per electron-volt ($\text{m}^{-3} \text{ eV}^{-1}$).

We can find an expression for the density of states by counting the number of standing electron matter waves that can fit into a box the size of the metal sample we are considering. This is analogous to counting the number of standing waves of sound that can exist in a closed organ pipe. Here the problem is three-dimensional (not one-dimensional) and the waves are matter waves (not sound waves). The result of such counting can be shown to be

$$N(E) = \frac{8\sqrt{2}\pi m^{3/2}}{h^3} E^{1/2} \quad (\text{density of states, } \text{m}^{-3} \text{ J}^{-1}), \quad (41-5)$$

where $m (= 9.109 \times 10^{-31} \text{ kg})$ is the electron mass, $h (= 6.626 \times 10^{-34} \text{ J} \cdot \text{s})$ is the Planck constant, E is the energy in joules at which $N(E)$ is to be evaluated, and $N(E)$ is in states per cubic meter per joule ($\text{m}^{-3} \text{ J}^{-1}$). To modify this equation so that the value of E is in electron-volts and the value of $N(E)$ is in states per cubic meter per electron-volt ($\text{m}^{-3} \text{ eV}^{-1}$), multiply the right side of the equation by $e^{3/2}$, where e is the fundamental charge, $1.602 \times 10^{-19} \text{ C}$. Figure 41-6 is a plot of such a modified version of Eq. 41-5. Note that nothing in Eq. 41-5 or Fig. 41-6 involves the shape, temperature, or composition of the sample.



CHECKPOINT 1

Is the spacing between adjacent energy levels at $E = 4 \text{ eV}$ in copper larger than, the same as, or smaller than the spacing at $E = 6 \text{ eV}$?

The density of energy levels increases upward in a band.

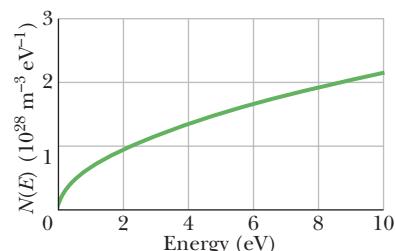


Fig. 41-6 The density of states $N(E)$ —that is, the number of electron energy levels per unit energy interval per unit volume—plotted as a function of electron energy. The density of states function simply counts the available states; it says nothing about whether these states are occupied by electrons.

Sample Problem

Number of states per electron volt in a metal

(a) Using the data of Fig. 41-6, determine the number of states per electron-volt at 7 eV in a metal sample with a volume V of $2 \times 10^{-9} \text{ m}^3$.

KEY IDEA

We can obtain the number of states per electron-volt at a given energy by using the density of states $N(E)$ at that energy and the sample's volume V .

Calculations: At an energy of 7 eV, we write

$$\left(\begin{array}{l} \text{number of states} \\ \text{per eV at 7 eV} \end{array} \right) = \left(\begin{array}{l} \text{density of states} \\ N(E) \text{ at 7 eV} \end{array} \right) \left(\begin{array}{l} \text{volume } V \\ \text{of sample} \end{array} \right).$$

From Fig. 41-6, we see that at an energy E of 7 eV, the density of states is about $1.8 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$. Thus,

$$\begin{aligned} \left(\begin{array}{l} \text{number of states} \\ \text{per eV at 7 eV} \end{array} \right) &= (1.8 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1})(2 \times 10^{-9} \text{ m}^3) \\ &= 3.6 \times 10^{19} \text{ eV}^{-1} \\ &\approx 4 \times 10^{19} \text{ eV}^{-1}. \end{aligned} \quad (\text{Answer})$$

(b) Next, determine the number of states N in the sample within a small energy range ΔE of 0.003 eV centered at 7 eV.

Calculation: From Eq. 41-5 and Fig. 41-6, we know that the density of states is a function of energy E . However, for an energy range ΔE that is small relative to E , we can approximate the density of states (and thus the number of states per electron-volt) to be constant. Thus, at an energy of 7 eV, we find the number of states N in the energy range ΔE of 0.003 eV as

$$\left(\begin{array}{l} \text{number of states } N \\ \text{in range } \Delta E \text{ at 7 eV} \end{array} \right) = \left(\begin{array}{l} \text{number of states} \\ \text{per eV at 7 eV} \end{array} \right) \left(\begin{array}{l} \text{energy} \\ \text{range } \Delta E \end{array} \right)$$

$$\begin{aligned} \text{or} \quad N &= (3.6 \times 10^{19} \text{ eV}^{-1})(0.003 \text{ eV}) \\ &= 1.1 \times 10^{17} \approx 1 \times 10^{17}. \end{aligned} \quad (\text{Answer})$$

(When you are asked for the number of states in a certain energy range, first see if that range is small enough to allow this type of approximation.)



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The Occupancy Probability $P(E)$

The ability of a metal to conduct electricity depends on the probability that available vacant levels will actually be occupied. Thus, another question arises: If an

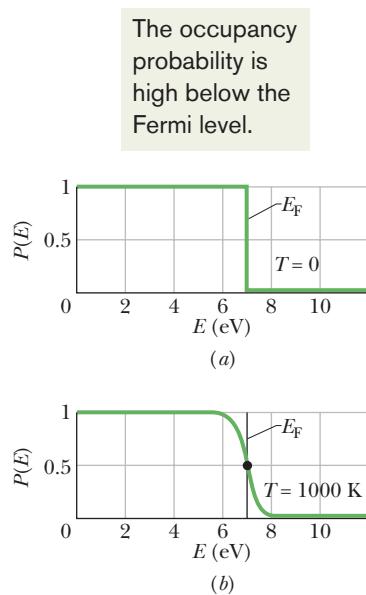


Fig. 41-7 The occupancy probability $P(E)$ is the probability that an energy level will be occupied by an electron. (a) At $T = 0$ K, $P(E)$ is unity for levels with energies E up to the Fermi energy E_F and zero for levels with higher energies. (b) At $T = 1000$ K, a few electrons whose energies were slightly less than the Fermi energy at $T = 0$ K move up to states with energies slightly greater than the Fermi energy. The dot on the curve shows that, for $E = E_F$, $P(E) = 0.5$.

energy level is available at energy E , what is the probability $P(E)$ that it is actually occupied by an electron? At $T = 0$ K, we know that for all levels with energies below the Fermi energy, $P(E) = 1$, corresponding to a certainty that the level is occupied. We also know that, at $T = 0$ K, for all levels with energies above the Fermi energy, $P(E) = 0$, corresponding to a certainty that the level is *not* occupied. Figure 41-7a illustrates this situation.

To find $P(E)$ at temperatures above absolute zero, we must use a set of quantum counting rules called **Fermi–Dirac statistics**, named for the physicists who introduced them. With these rules, the **occupancy probability** $P(E)$ is

$$P(E) = \frac{1}{e^{(E-E_F)/kT} + 1} \quad (\text{occupancy probability}), \quad (41-6)$$

in which E_F is the Fermi energy. Note that $P(E)$ depends not on the energy E of the level but only on the difference $E - E_F$, which may be positive or negative.

To see whether Eq. 41-6 describes Fig. 41-7a, we substitute $T = 0$ K in it. Then, for $E < E_F$, the exponential term in Eq. 41-6 is $e^{-\infty}$, or zero; so $P(E) = 1$, in agreement with Fig. 41-7a.

For $E > E_F$, the exponential term is $e^{+\infty}$; so $P(E) = 0$, again in agreement with Fig. 41-7a.

Figure 41-7b is a plot of $P(E)$ for $T = 1000$ K. It shows that, as stated above, changes in the distribution of electrons among the available states involve only states whose energies are near the Fermi energy E_F . Note that if $E = E_F$ (no matter what the temperature T), the exponential term in Eq. 41-6 is $e^0 = 1$ and $P(E) = 0.5$. This leads us to a more useful definition of the Fermi energy:



The Fermi energy of a given material is the energy of a quantum state that has the probability 0.5 of being occupied by an electron.

Figures 41-7a and b are plotted for copper, which has a Fermi energy of 7.0 eV. Thus, for copper both at $T = 0$ K and at $T = 1000$ K, a state at energy $E = 7.0$ eV has a probability of 0.5 of being occupied.

Sample Problem

Probability of occupancy of an energy state in a metal

- (a) What is the probability that a quantum state whose energy is 0.10 eV above the Fermi energy will be occupied? Assume a sample temperature of 800 K.

KEY IDEA

The occupancy probability of any state in a metal can be found from Fermi–Dirac statistics according to Eq. 41-6.

Calculations: Let's start with the exponent in Eq. 41-6:

$$\frac{E - E_F}{kT} = \frac{0.10 \text{ eV}}{(8.62 \times 10^{-5} \text{ eV/K})(800 \text{ K})} = 1.45.$$

Inserting this exponent into Eq. 41-6 yields

$$P(E) = \frac{1}{e^{1.45} + 1} = 0.19 \text{ or } 19\%. \quad (\text{Answer})$$

- (b) What is the probability of occupancy for a state that is 0.10 eV *below* the Fermi energy?

Calculation: The Key Idea of part (a) applies here also except that now the state has an energy *below* the Fermi energy. Thus, the exponent in Eq. 41-6 has the same magnitude we found in part (a) but is negative, and that makes the denominator smaller. Equation 41-6 now yields

$$P(E) = \frac{1}{e^{-1.45} + 1} = 0.81 \text{ or } 81\%. \quad (\text{Answer})$$

For states below the Fermi energy, we are often more interested in the probability that the state is *not* occupied. This probability is just $1 - P(E)$, or 19%. Note that it is the same as the probability of occupancy in (a).



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How Many Occupied States Are There?

Equation 41-5 and Fig. 41-6 tell us how the available states are distributed in energy. The occupancy probability of Eq. 41-6 gives us the probability that any given state will actually be occupied by an electron. To find $N_o(E)$, the **density of occupied states**, we must multiply each available state by the corresponding value of the occupancy probability; that is,

$$\left(\begin{array}{c} \text{density of occupied states} \\ N_o(E) \text{ at energy } E \end{array} \right) = \left(\begin{array}{c} \text{density of states} \\ N(E) \text{ at energy } E \end{array} \right) \left(\begin{array}{c} \text{occupancy probability} \\ P(E) \text{ at energy } E \end{array} \right)$$

or

$$N_o(E) = N(E) P(E) \quad (\text{density of occupied states}). \quad (41-7)$$

Figure 41-8a is a plot of Eq. 41-7 for copper at $T = 0$ K. It is found by multiplying, at each energy, the value of the density of states function (Fig. 41-6) by the value of the occupancy probability for absolute zero (Fig. 41-7a). Figure 41-8b, calculated similarly, shows the density of occupied states for copper at $T = 1000$ K.

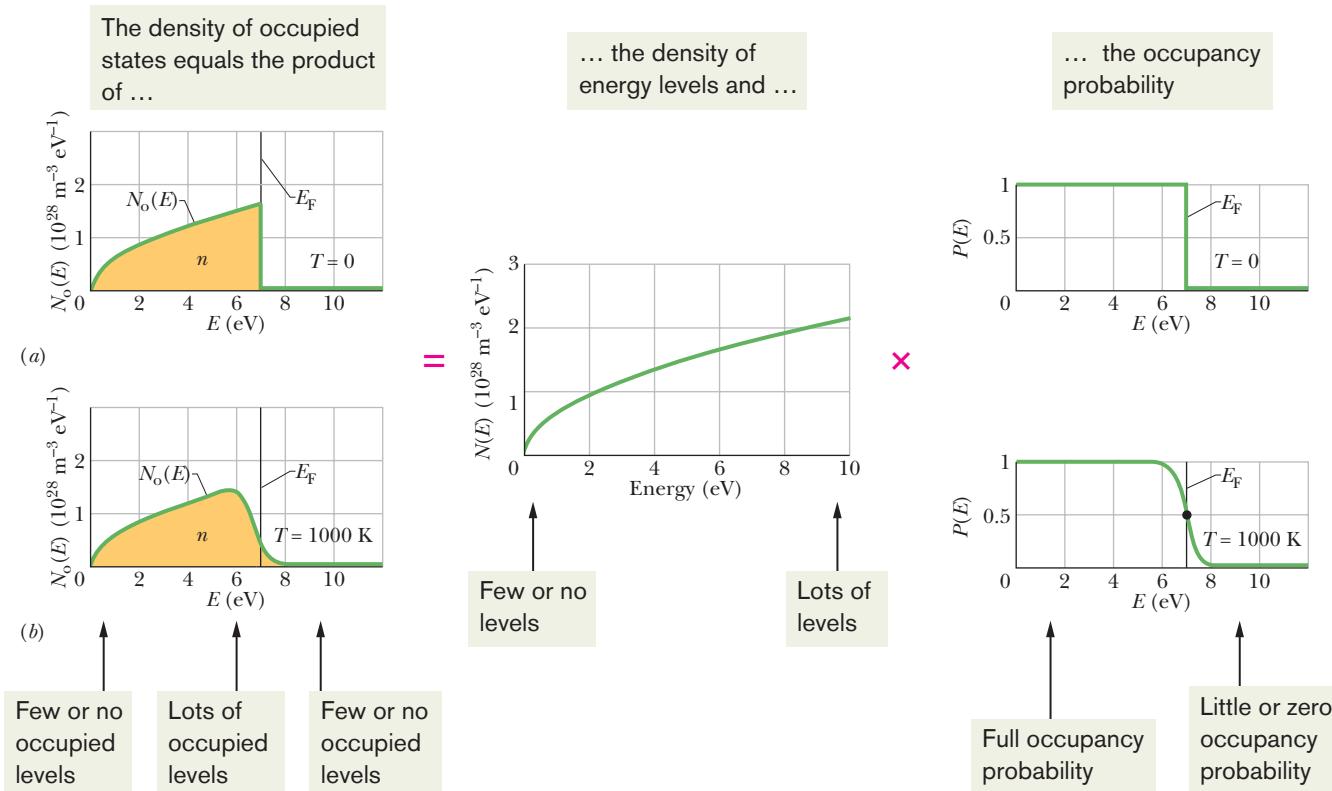


Fig. 41-8 (a) The density of occupied states $N_o(E)$ for copper at absolute zero. The area under the curve is the number density of electrons n . Note that all states with energies up to the Fermi energy $E_F = 7$ eV are occupied, and all those with energies above the Fermi energy are vacant. (b) The same for copper at $T = 1000$ K. Note that only electrons with energies near the Fermi energy have been affected and redistributed.

Sample Problem**Number of occupied states in an energy range in a metal**

A lump of copper (Fermi energy = 7.0 eV) has volume $2 \times 10^{-9} \text{ m}^3$. How many occupied states per eV lie in a narrow energy range around 7.0 eV?

KEY IDEAS

(1) First we want the density of occupied states $N_o(E)$ as given by Eq. 41-7 ($N_o(E) = N(E) P(E)$). (2) Because we want to evaluate quantities for a narrow energy range around 7.0 eV (the Fermi energy for copper), the occupancy probability $P(E)$ is 0.50.

Calculations: From Fig. 41-6, we see that the density of states at 7 eV is about $1.8 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$. Thus, Eq. 41-7 tells us that the density of occupied states is

$$\begin{aligned} N_o(E) &= N(E) P(E) = (1.8 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1})(0.50) \\ &= 0.9 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}. \end{aligned}$$

Next, we write

$$\begin{aligned} \left(\begin{array}{l} \text{number of occupied} \\ \text{states per eV at 7 eV} \end{array} \right) &= \left(\begin{array}{l} \text{density of occupied} \\ \text{states } N_o(E) \text{ at 7 eV} \end{array} \right) \\ &\quad \times \left(\begin{array}{l} \text{volume } V \\ \text{of sample} \end{array} \right). \end{aligned}$$

Substituting for $N_o(E)$ and V gives us

$$\begin{aligned} \left(\begin{array}{l} \text{number of occupied} \\ \text{states per eV} \\ \text{at 7 eV} \end{array} \right) &= (0.9 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1})(2 \times 10^{-9} \text{ m}^3) \\ &= 1.8 \times 10^{19} \text{ eV}^{-1} \\ &\approx 2 \times 10^{19} \text{ eV}^{-1}. \end{aligned} \quad (\text{Answer})$$



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Calculating the Fermi Energy

Suppose we add up (via integration) the number of occupied states per unit volume in Fig. 41-8a (for $T = 0 \text{ K}$) at all energies between $E = 0$ and $E = E_F$. The result must equal n , the number of conduction electrons per unit volume for the metal. In equation form, we have

$$n = \int_0^{E_F} N_o(E) dE. \quad (41-8)$$

(Graphically, the integral here represents the area under the distribution curve of Fig. 41-8a.) Because $P(E) = 1$ for all energies below the Fermi energy when $T = 0 \text{ K}$, Eq. 41-7 tells us we can replace $N_o(E)$ in Eq. 41-8 with $N(E)$ and then use Eq. 41-8 to find the Fermi energy E_F . If we substitute Eq. 41-5 into Eq. 41-8, we find that

$$n = \frac{8\sqrt{2}\pi m^{3/2}}{h^3} \int_0^{E_F} E^{1/2} dE = \frac{8\sqrt{2}\pi m^{3/2}}{h^3} \frac{2E_F^{3/2}}{3},$$

in which m is the electron mass. Solving for E_F now leads to

$$E_F = \left(\frac{3}{16\sqrt{2}\pi} \right)^{2/3} \frac{h^2}{m} n^{2/3} = \frac{0.121h^2}{m} n^{2/3}. \quad (41-9)$$

Thus, when we know n , the number of conduction electrons per unit volume for a metal, we can find the Fermi energy for that metal.

41-6 Semiconductors

If you compare Fig. 41-9a with Fig. 41-4, you can see that the band structure of a semiconductor is like that of an insulator. The main difference is that the semiconductor has a much smaller energy gap E_g between the top of the highest filled

band (called the **valence band**) and the bottom of the vacant band just above it (called the **conduction band**). Thus, there is no doubt that silicon ($E_g = 1.1$ eV) is a semiconductor and diamond ($E_g = 5.5$ eV) is an insulator. In silicon—but not in diamond—there is a real possibility that thermal agitation at room temperature will cause electrons to jump the gap from valence to conduction band.

In Table 41-1 we compared three basic electrical properties of copper, our prototype metallic conductor, and silicon, our prototype semiconductor. Let us look again at that table, one row at a time, to see how a semiconductor differs from a metal.

Number Density of Charge Carriers n

The bottom row of Table 41-1 shows that copper has far more charge carriers per unit volume than silicon, by a factor of about 10^{13} . For copper, each atom contributes one electron, its single valence electron, to the conduction process. Charge carriers in silicon arise only because, at thermal equilibrium, thermal agitation causes a certain (very small) number of valence-band electrons to jump the energy gap into the conduction band, leaving an equal number of unoccupied energy states, called **holes**, in the valence band. Figure 41-9b shows the situation.

Both the electrons in the conduction band and the holes in the valence band serve as charge carriers. The holes do so by permitting a certain freedom of movement to the electrons remaining in the valence band, electrons that, in the absence of holes, would be gridlocked. If an electric field \vec{E} is set up in a semiconductor, the electrons in the valence band, being negatively charged, tend to drift in the direction opposite \vec{E} . This causes the positions of the holes to drift in the direction of \vec{E} . In effect, the holes behave like moving particles of charge $+e$.

It may help to think of a row of cars parked bumper to bumper, with the lead car at one car's length from a barrier and the empty one-car-length distance being an available parking space. If the leading car moves forward to the barrier, it opens up a parking space behind it. The second car can then move up to fill that space, allowing the third car to move up, and so on. The motions of the many cars toward the barrier are most simply analyzed by focusing attention on the drift of the single “hole” (parking space) away from the barrier.

In semiconductors, conduction by holes is just as important as conduction by electrons. In thinking about hole conduction, we can assume that all unoccupied states in the valence band are occupied by particles of charge $+e$ and that all electrons in the valence band have been removed, so that these positive charge carriers can move freely throughout the band.

Resistivity ρ

Recall from Chapter 26 that the resistivity ρ of a material is $m/e^2 n \tau$, where m is the electron mass, e is the fundamental charge, n is the number of charge carriers per unit volume, and τ is the mean time between collisions of the charge carriers. Table 41-1 shows that, at room temperature, the resistivity of silicon is higher than that of copper by a factor of about 10^{11} . This vast difference can be accounted for by the vast difference in n . Other factors enter, but their effect on the resistivity is swamped by the enormous difference in n .

Temperature Coefficient of Resistivity α

Recall that α (see Eq. 26-17) is the fractional change in resistivity per unit change in temperature:

$$\alpha = \frac{1}{\rho} \frac{d\rho}{dT}. \quad (41-10)$$

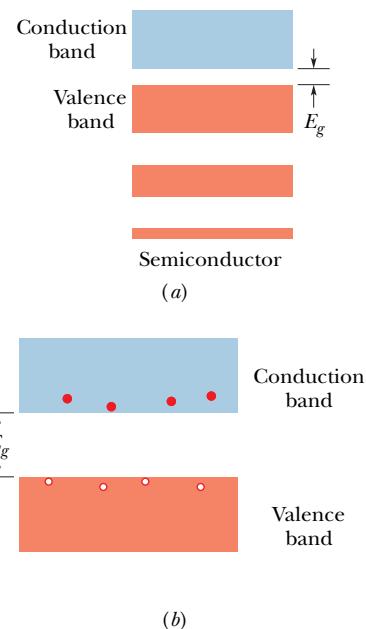


Fig. 41-9 (a) The band-gap pattern for a semiconductor. It resembles that of an insulator (see Fig. 41-4) except that here the energy gap E_g is much smaller; thus electrons, because of their thermal agitation, have some reasonable probability of being able to jump the gap. (b) Thermal agitation has caused a few electrons to jump the gap from the valence band to the conduction band, leaving an equal number of holes in the valence band.

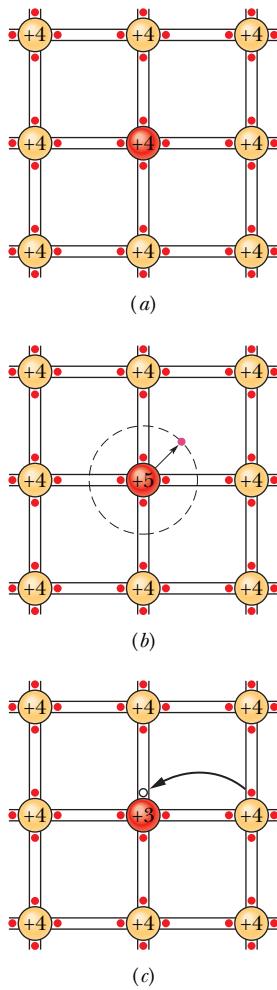


Fig. 41-10 (a) A flattened-out representation of the lattice structure of pure silicon. Each silicon ion is coupled to its four nearest neighbors by a two-electron covalent bond (represented by a pair of red dots between two parallel black lines). The electrons belong to the bond—not to the individual atoms—and form the valence band of the sample. (b) One silicon atom is replaced by a phosphorus atom (valence = 5). The “extra” electron is only loosely bound to its ion core and may easily be elevated to the conduction band, where it is free to wander through the volume of the lattice. (c) One silicon atom is replaced by an aluminum atom (valence = 3). There is now a hole in one of the covalent bonds and thus in the valence band of the sample. The hole can easily migrate through the lattice as electrons from neighboring bonds move in to fill it. Here the hole migrates rightward.

The resistivity of copper *increases* with temperature (that is, $d\rho/dT > 0$) because collisions of copper’s charge carriers occur more frequently at higher temperatures. Thus, α is *positive* for copper.

The collision frequency also increases with temperature for silicon. However, the resistivity of silicon actually *decreases* with temperature ($d\rho/dT < 0$) because the number of charge carriers n (electrons in the conduction band and holes in the valence band) increases so rapidly with temperature. (More electrons jump the gap from the valence band to the conduction band.) Thus, the fractional change α is *negative* for silicon.

41-7 Doped Semiconductors

The usefulness of semiconductors in technology can be greatly improved by introducing a small number of suitable replacement atoms (called **impurities**) into the semiconductor lattice—a process called **doping**. Typically, only about 1 silicon atom in 10^7 is replaced by a dopant atom in the doped semiconductor. Essentially all modern semiconducting devices are based on doped material. Such materials are of two types, called ***n*-type** and ***p*-type**; we discuss each in turn.

n-Type Semiconductors

The electrons in an isolated silicon atom are arranged in subshells according to the scheme

$$1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^2,$$

in which, as usual, the superscripts (which add to 14, the atomic number of silicon) represent the numbers of electrons in the specified subshells.

Figure 41-10a is a flattened-out representation of a portion of the lattice of pure silicon in which the portion has been projected onto a plane; compare the figure with Fig. 41-1b, which represents the unit cell of the lattice in three dimensions. Each silicon atom contributes its pair of $3s$ electrons and its pair of $3p$ electrons to form a rigid two-electron covalent bond with each of its four nearest neighbors. (A covalent bond is a link between two atoms in which the atoms share a pair of electrons.) The four atoms that lie within the unit cell in Fig. 41-1b show these four bonds.

The electrons that form the silicon–silicon bonds constitute the valence band of the silicon sample. If an electron is torn from one of these bonds so that it becomes free to wander throughout the lattice, we say that the electron has been raised from the valence band to the conduction band. The minimum energy required to do this is the gap energy E_g .

Because four of its electrons are involved in bonds, each silicon “atom” is actually an ion consisting of an inert neon-like electron cloud (containing 10 electrons) surrounding a nucleus whose charge is $+14e$, where 14 is the atomic number of silicon. The net charge of each of these ions is thus $+4e$, and the ions are said to have a *valence number* of 4.

In Fig. 41-10b the central silicon ion has been replaced by an atom of phosphorus (valence = 5). Four of the valence electrons of the phosphorus form bonds with the four surrounding silicon ions. The fifth (“extra”) electron is only loosely bound to the phosphorus ion core. On an energy-band diagram, we usually say that such an electron occupies a localized energy state that lies within the energy gap, at an average energy interval E_d below the bottom of the conduction band; this is indicated in Fig. 41-11a. Because $E_d \ll E_g$, the energy required to excite electrons from these levels into the conduction band is much less than that required to excite silicon valence electrons into the conduction band.

The phosphorus atom is called a **donor** atom because it readily *donates* an electron to the conduction band. In fact, at room temperature virtually *all* the

41-7 DOPED SEMICONDUCTORS

1153

electrons contributed by the donor atoms are in the conduction band. By adding donor atoms, it is possible to increase greatly the number of electrons in the conduction band, by a factor very much larger than Fig. 41-11a suggests.

Semiconductors doped with donor atoms are called ***n*-type semiconductors**; the *n* stands for *negative*, to imply that the negative charge carriers introduced into the conduction band greatly outnumber the positive charge carriers, which are the holes in the valence band. In *n*-type semiconductors, the electrons are called the **majority carriers** and the holes are called the **minority carriers**.

p-Type Semiconductors

Now consider Fig. 41-10c, in which one of the silicon atoms (valence = 4) has been replaced by an atom of aluminum (valence = 3). The aluminum atom can bond covalently with only three silicon atoms, and so there is now a “missing” electron (a hole) in one aluminum–silicon bond. With a small expenditure of energy, an electron can be torn from a neighboring silicon–silicon bond to fill this hole, thereby creating a hole in *that* bond. Similarly, an electron from some other bond can be moved to fill the newly created hole. In this way, the hole can migrate through the lattice.

The aluminum atom is called an **acceptor** atom because it readily *accepts* an electron from a neighboring bond—that is, from the valence band of silicon. As Fig. 41-11b suggests, this electron occupies a localized energy state that lies within the energy gap, at an average energy interval E_a above the top of the valence band. Because this energy interval E_a is small, valence electrons are easily bumped up to the acceptor level, leaving holes in the valence band. Thus, by adding acceptor atoms, it is possible to greatly increase the number of holes in the valence band, by a factor much larger than Fig. 41-11b suggests. In silicon at room temperature, virtually *all* the acceptor levels are occupied by electrons.

Semiconductors doped with acceptor atoms are called ***p*-type semiconductors**; the *p* stands for *positive* to imply that the holes introduced into the valence band, which behave like positive charge carriers, greatly outnumber the electrons in the conduction band. In *p*-type semiconductors, holes are the majority carriers and electrons are the minority carriers.

Table 41-2 summarizes the properties of a typical *n*-type and a typical *p*-type semiconductor. Note particularly that the donor and acceptor ion cores, although they are charged, are not charge *carriers* because they are fixed in place.

Table 41-2
Properties of Two Doped Semiconductors

Property	Type of Semiconductor	
	<i>n</i>	<i>p</i>
Matrix material	Silicon	Silicon
Matrix nuclear charge	+14e	+14e
Matrix energy gap	1.2 eV	1.2 eV
Dopant	Phosphorus	Aluminum
Type of dopant	Donor	Acceptor
Majority carriers	Electrons	Holes
Minority carriers	Holes	Electrons
Dopant energy gap	$E_d = 0.045 \text{ eV}$	$E_a = 0.067 \text{ eV}$
Dopant valence	5	3
Dopant nuclear charge	+15e	+13e
Dopant net ion charge	+e	-e

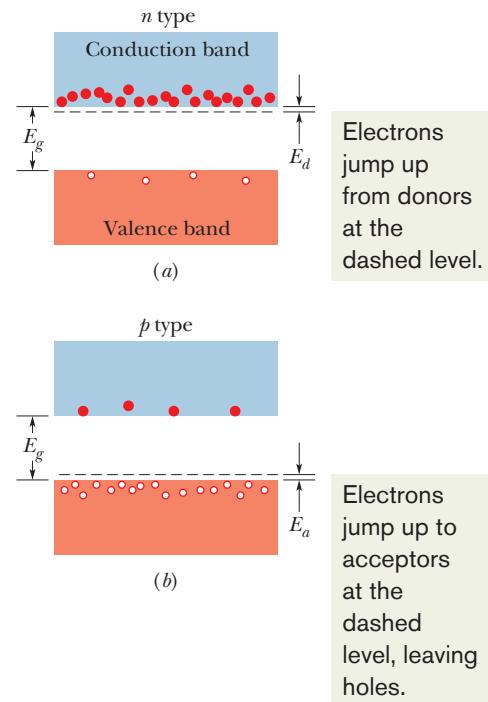


Fig. 41-11 (a) In a doped *n*-type semiconductor, the energy levels of donor electrons lie a small interval E_d below the bottom of the conduction band. Because donor electrons can be easily excited to the conduction band, there are now many more electrons in that band. The valence band contains the same small number of holes as before the dopant was added. (b) In a doped *p*-type semiconductor, the acceptor levels lie a small energy interval E_a above the top of the valence band. There are now many more holes in the valence band. The conduction band contains the same small number of electrons as before the dopant was added. The ratio of majority carriers to minority carriers in both (a) and (b) is very much greater than is suggested by these diagrams.

Sample Problem**Doping silicon with phosphorus**

The number density n_0 of conduction electrons in pure silicon at room temperature is about 10^{16} m^{-3} . Assume that, by doping the silicon lattice with phosphorus, we want to increase this number by a factor of a million (10^6). What fraction of silicon atoms must we replace with phosphorus atoms? (Recall that at room temperature, thermal agitation is so effective that essentially every phosphorus atom donates its “extra” electron to the conduction band.)

Number of phosphorus atoms: Because each phosphorus atom contributes one conduction electron and because we want the total number density of conduction electrons to be $10^6 n_0$, the number density of phosphorus atoms n_p must be given by

$$10^6 n_0 = n_0 + n_p.$$

$$\begin{aligned} \text{Then } n_p &= 10^6 n_0 - n_0 \approx 10^6 n_0 \\ &= (10^6)(10^{16} \text{ m}^{-3}) = 10^{22} \text{ m}^{-3}. \end{aligned}$$

This tells us that we must add 10^{22} atoms of phosphorus per cubic meter of silicon.

Fraction of silicon atoms: We can find the number density n_{Si} of silicon atoms in pure silicon (before the doping) from Eq. 41-4, which we can write as

$$\begin{aligned} (\text{number of atoms}) \\ \text{in sample} \\ = \frac{(\text{silicon density})(\text{sample volume } V)}{(\text{silicon molar mass } M_{\text{Si}})/N_A}. \end{aligned}$$



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Dividing both sides by the sample volume V to get the number density of silicon atoms n_{Si} on the left, we then have

$$n_{\text{Si}} = \frac{(\text{silicon density})N_A}{M_{\text{Si}}}.$$

Appendix F tells us that the density of silicon is 2.33 g/cm^3 ($= 2330 \text{ kg/m}^3$) and the molar mass of silicon is 28.1 g/mol ($= 0.0281 \text{ kg/mol}$). Thus, we have

$$\begin{aligned} n_{\text{Si}} &= \frac{(2330 \text{ kg/m}^3)(6.02 \times 10^{23} \text{ atoms/mol})}{0.0281 \text{ kg/mol}} \\ &= 5 \times 10^{28} \text{ atoms/m}^3 = 5 \times 10^{28} \text{ m}^{-3}. \end{aligned}$$

The fraction we seek is approximately

$$\frac{n_p}{n_{\text{Si}}} = \frac{10^{22} \text{ m}^{-3}}{5 \times 10^{28} \text{ m}^{-3}} = \frac{1}{5 \times 10^6}. \quad (\text{Answer})$$

If we replace only *one silicon atom in five million* with a phosphorus atom, the number of electrons in the conduction band will be increased by a factor of a million.

How can such a tiny admixture of phosphorus have what seems to be such a big effect? The answer is that, although the effect is very significant, it is not “big.” The number density of conduction electrons was 10^{16} m^{-3} before doping and 10^{22} m^{-3} after doping. For copper, however, the conduction-electron number density (given in Table 41-1) is about 10^{29} m^{-3} . Thus, even after doping, the number density of conduction electrons in silicon remains much less than that of a typical metal, such as copper, by a factor of about 10^7 .

41-8 The p-n Junction

A **p-n junction** (Fig. 41-12a), essential to most semiconductor devices, is a single semiconductor crystal that has been selectively doped so that one region is *n*-type material and the adjacent region is *p*-type material. Let’s assume that the junction has been formed mechanically by jamming together a bar of *n*-type semiconductor and a bar of *p*-type semiconductor. Thus, the transition from one region to the other is perfectly sharp, occurring at a single **junction plane**.

Let us discuss the motions of electrons and holes just after the *n*-type bar and the *p*-type bar, both electrically neutral, have been jammed together to form the junction. We first examine the majority carriers, which are electrons in the *n*-type material and holes in the *p*-type material.

Motions of the Majority Carriers

If you burst a helium-filled balloon, helium atoms will diffuse (spread) outward into the surrounding air. This happens because there are very few helium atoms

in normal air. In more formal language, there is a helium *density gradient* at the balloon–air interface (the number density of helium atoms varies across the interface); the helium atoms move so as to reduce the gradient.

In the same way, electrons on the *n* side of Fig. 41-12a that are close to the junction plane tend to diffuse across it (from right to left in the figure) and into the *p* side, where there are very few free electrons. Similarly, holes on the *p* side that are close to the junction plane tend to diffuse across that plane (from left to right) and into the *n* side, where there are very few holes. The motions of both the electrons and the holes contribute to a **diffusion current** I_{diff} , conventionally directed from left to right as indicated in Fig. 41-12d.

Recall that the *n*-side is studded throughout with positively charged donor ions, fixed firmly in their lattice sites. Normally, the excess positive charge of each of these ions is compensated electrically by one of the conduction-band electrons. When an *n*-side electron diffuses across the junction plane, however, the diffusion “uncovers” one of these donor ions, thus introducing a fixed positive charge near the junction plane on the *n* side. When the diffusing electron arrives on the *p* side, it quickly combines with an acceptor ion (which lacks one electron), thus introducing a fixed negative charge near the junction plane on the *p* side.

In this way electrons diffusing through the junction plane from right to left in Fig. 41-12a result in a buildup of **space charge** on each side of the junction plane, as indicated in Fig. 41-12b. Holes diffusing through the junction plane from left to right have exactly the same effect. (Take the time now to convince yourself of that.) The motions of both majority carriers—electrons and holes—contribute to the buildup of these two space charge regions, one positive and one negative. These two regions form a **depletion zone**, so named because it is relatively free of mobile charge carriers; its width is shown as d_0 in Fig. 41-12b.

The buildup of space charge generates an associated **contact potential difference** V_0 across the depletion zone, as Fig. 41-12c shows. This potential difference limits further diffusion of electrons and holes across the junction plane. Negative charges tend to avoid regions of low potential. Thus, an electron approaching the junction plane from the right in Fig. 41-12b is moving toward a region of low potential and would tend to turn back into the *n* side. Similarly, a positive charge (a hole) approaching the junction plane from the left is moving toward a region of high potential and would tend to turn back into the *p* side.

Motions of the Minority Carriers

As Fig. 41-11a shows, although the majority carriers in *n*-type material are electrons, there are a few holes. Likewise in *p*-type material (Fig. 41-11b), although the majority carriers are holes, there are also a few electrons. These few holes and electrons are the minority carriers in the corresponding materials.

Although the potential difference V_0 in Fig. 41-12c acts as a barrier for the majority carriers, it is a downhill trip for the minority carriers, be they electrons on the *p* side or holes on the *n* side. Positive charges (holes) tend to seek regions of low potential; negative charges (electrons) tend to seek regions of high potential. Thus, both types of minority carriers are *swept across* the junction plane by the contact potential difference and together constitute a **drift current** I_{drift} across the junction plane from right to left, as Fig. 41-12d indicates.

Thus, an isolated *p-n* junction is in an equilibrium state in which a contact potential difference V_0 exists between its ends. At equilibrium, the average diffusion current I_{diff} that moves through the junction plane from the *p* side to the *n* side is just balanced by an average drift current I_{drift} that moves in the opposite direction. These two currents cancel because the net current through the junction plane must be zero; otherwise charge would be transferred without limit from one end of the junction to the other.

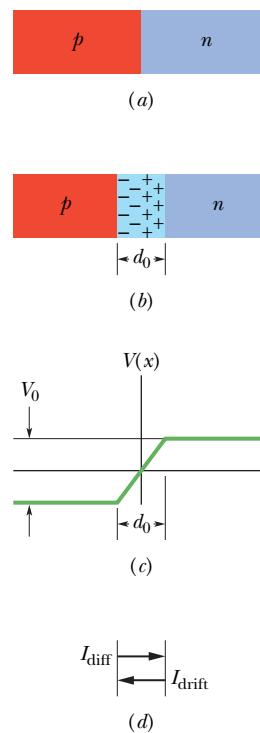


Fig. 41-12 (a) A *p-n* junction. (b) Motions of the majority charge carriers across the junction plane uncover a space charge associated with uncompensated donor ions (to the right of the plane) and acceptor ions (to the left). (c) Associated with the space charge is a contact potential difference V_0 across d_0 . (d) The diffusion of majority carriers (both electrons and holes) across the junction plane produces a diffusion current I_{diff} . (In a real *p-n* junction, the boundaries of the depletion zone would not be sharp, as shown here, and the contact potential curve (c) would be smooth, with no sharp corners.)

 **CHECKPOINT 2**

Which of the following five currents across the junction plane of Fig. 41-12a must be zero?

- the net current due to holes, both majority and minority carriers included
- the net current due to electrons, both majority and minority carriers included
- the net current due to both holes and electrons, both majority and minority carriers included
- the net current due to majority carriers, both holes and electrons included
- the net current due to minority carriers, both holes and electrons included

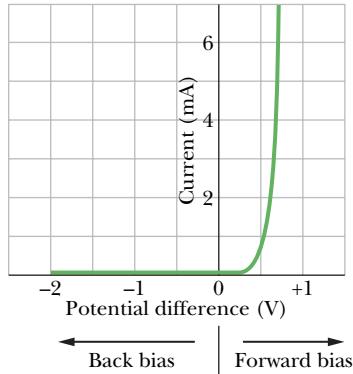


Fig. 41-13 A current–voltage plot for a *p*–*n* junction, showing that the junction is highly conducting when forward-biased and essentially nonconducting when back-biased.

41-9 The Junction Rectifier

Look now at Fig. 41-13. It shows that, if we place a potential difference across a *p*–*n* junction in one direction (here labeled + and “Forward bias”), there will be a current through the junction. However, if we reverse the direction of the potential difference, there will be approximately zero current through the junction.

One application of this property is the **junction rectifier**, whose symbol is shown in Fig. 41-14*b*; the arrowhead corresponds to the *p*-type end of the device and points in the allowed direction of conventional current. A sine wave input potential to the device (Fig. 41-14*a*) is transformed to a half-wave output potential (Fig. 41-14*c*) by the junction rectifier; that is, the rectifier acts as essentially a closed switch (zero resistance) for one polarity of the input potential and as essentially an open switch (infinite resistance) for the other. The average input voltage is zero, but the average output voltage is not. Thus, a junction rectifier can be used as part of an apparatus to convert an alternating potential difference into a constant potential difference, as for a power supply.

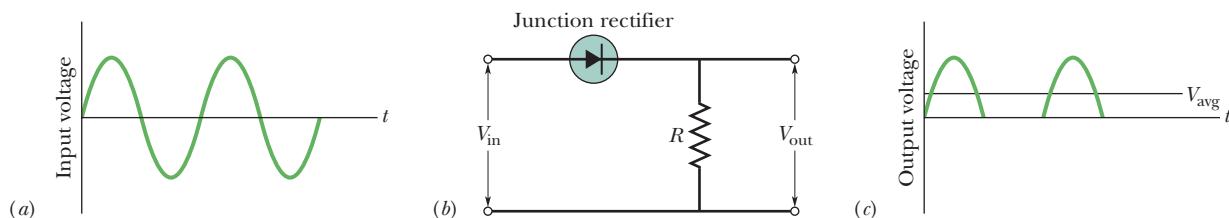


Fig. 41-14 A *p*–*n* junction connected as a junction rectifier. The action of the circuit in (b) is to pass the positive half of the input wave form in (a) but to suppress the negative half. The average potential of the input wave form is zero; that of the output wave form in (c) has a positive value V_{avg} .

Figure 41-15 shows why a *p*–*n* junction operates as a junction rectifier. In Fig. 41-15*a*, a battery is connected across the junction with its positive terminal connected at the *p* side. In this **forward-bias connection**, the *p* side becomes more positive and the *n* side becomes more negative, thus *decreasing* the height of

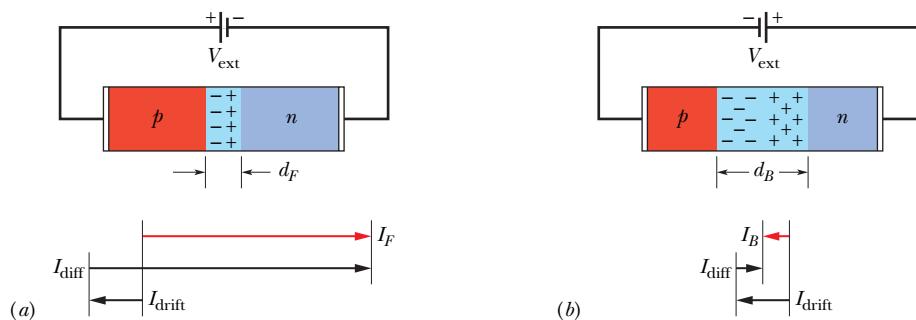


Fig. 41-15 (a) The forward-bias connection of a *p*–*n* junction, showing the narrowed depletion zone and the large forward current I_F . (b) The back-bias connection, showing the widened depletion zone and the small back current I_B .

the potential barrier V_0 of Fig. 41-12c. More of the majority carriers can now surmount this smaller barrier; hence, the diffusion current I_{diff} increases markedly.

The minority carriers that form the drift current, however, sense no barrier; so the drift current I_{drift} is not affected by the external battery. The nice current balance that existed at zero bias (see Fig. 41-12d) is thus upset, and, as shown in Fig. 41-15a, a large net forward current I_F appears in the circuit.

Another effect of forward bias is to narrow the depletion zone, as a comparison of Fig. 41-12b and Fig. 41-15a shows. The depletion zone narrows because the reduced potential barrier associated with forward bias must be associated with a smaller space charge. Because the ions producing the space charge are fixed in their lattice sites, a reduction in their number can come about only through a reduction in the width of the depletion zone.

Because the depletion zone normally contains very few charge carriers, it is normally a region of high resistivity. However, when its width is substantially reduced by a forward bias, its resistance is also reduced substantially, as is consistent with the large forward current.

Figure 41-15b shows the **back-bias** connection, in which the negative terminal of the battery is connected at the *p*-type end of the *p-n* junction. Now the applied emf *increases* the contact potential difference, the diffusion current *decreases* substantially while the drift current remains unchanged, and a relatively *small* back current I_B results. The depletion zone *widens*, its *high* resistance being consistent with the *small* back current I_B .

41-10 The Light-Emitting Diode (LED)

Nowadays, we can hardly avoid the brightly colored “electronic” numbers that glow at us from cash registers and gasoline pumps, microwave ovens and alarm clocks, and we cannot seem to do without the invisible infrared beams that control elevator doors and operate television sets via remote control. In nearly all cases this light is emitted from a *p-n* junction operating as a **light-emitting diode** (LED). How can a *p-n* junction generate light?

Consider first a simple semiconductor. When an electron from the bottom of the conduction band falls into a hole at the top of the valence band, an energy E_g equal to the gap width is released. In silicon, germanium, and many other semiconductors, this energy is largely transformed into thermal energy of the vibrating lattice, and as a result, no light is emitted.

In some semiconductors, however, including gallium arsenide, the energy can be emitted as a photon of energy hf at wavelength

$$\lambda = \frac{c}{f} = \frac{c}{E_g/h} = \frac{hc}{E_g}. \quad (41-11)$$

To emit enough light to be useful as an LED, the material must have a suitably large number of electron-hole transitions. This condition is not satisfied by a pure semiconductor because, at room temperature, there are simply not enough electron-hole pairs. As Fig. 41-11 suggests, doping will not help. In doped *n*-type material the number of conduction electrons is greatly increased, but there are not enough holes for them to combine with; in doped *p*-type material there are plenty of holes but not enough electrons to combine with them. Thus, neither a pure semiconductor nor a doped semiconductor can provide enough electron-hole transitions to serve as a practical LED.

What we need is a semiconductor material with a very large number of electrons in the conduction band *and* a correspondingly large number of holes in the valence band. A device with this property can be fabricated by placing a strong forward bias on a heavily doped *p-n* junction, as in Fig. 41-16. In such an arrangement the current I through the device serves to inject electrons into the *n*-type

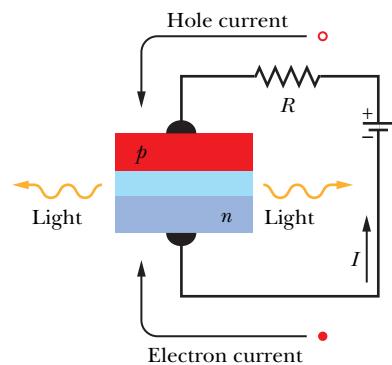


Fig. 41-16 A forward-biased *p-n* junction, showing electrons being injected into the *n*-type material and holes into the *p*-type material. (Holes move in the conventional direction of the current I , equivalent to electrons moving in the opposite direction.) Light is emitted from the narrow depletion zone each time an electron and a hole combine across that zone.

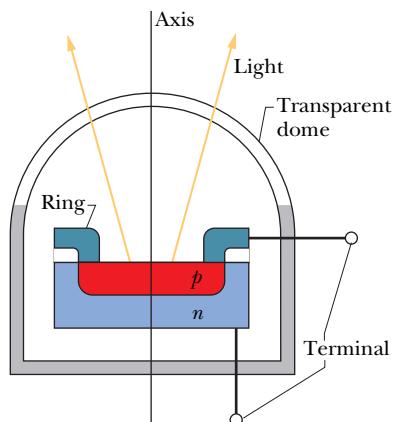


Fig. 41-17 Cross section of an LED (the device has rotational symmetry about the central axis). The *p*-type material, which is thin enough to transmit light, is in the form of a circular disk. A connection is made to the *p*-type material through a circular metal ring that touches the disk at its periphery. The depletion zone between the *n*-type material and the *p*-type material is not shown.

material and to inject holes into the *p*-type material. If the doping is heavy enough and the current is great enough, the depletion zone can become very narrow, perhaps only a few micrometers wide. The result is a great number density of electrons in the *n*-type material facing a correspondingly great number density of holes in the *p*-type material, across the narrow depletion zone. With such great number densities so near each other, many electron–hole combinations occur, causing light to be emitted from that zone. Figure 41-17 shows the construction of an actual LED.

Commercial LEDs designed for the visible region are commonly based on gallium suitably doped with arsenic and phosphorus atoms. An arrangement in which 60% of the nongallium sites are occupied by arsenic ions and 40% by phosphorus ions results in a gap width E_g of about 1.8 eV, corresponding to red light. Other doping and transition-level arrangements make it possible to construct LEDs that emit light in essentially any desired region of the visible and near-visible spectra.

The Photodiode

Passing a current through a suitably arranged *p-n* junction can generate light. The reverse is also true; that is, shining light on a suitably arranged *p-n* junction can produce a current in a circuit that includes the junction. This is the basis for the **photodiode**.

When you click your television remote control, an LED in the device sends out a coded sequence of pulses of infrared light. The receiving device in your television set is an elaboration of the simple (two-terminal) photodiode that not only detects the infrared signals but also amplifies them and transforms them into electrical signals that change the channel or adjust the volume, among other tasks.

The Junction Laser

In the arrangement of Fig. 41-16, there are many electrons in the conduction band of the *n*-type material and many holes in the valence band of the *p*-type material. Thus, there is a **population inversion** for the electrons; that is, there are more electrons in higher energy levels than in lower energy levels. As we discussed in Section 40-12, this is normally a necessary—but not a sufficient—condition for laser action.

When a single electron moves from the conduction band to the valence band, it can release its energy as a photon. This photon can stimulate a second electron to fall into the valence band, producing a second photon by stimulated emission. In this way, if the current through the junction is great enough, a chain reaction of stimulated emission events can occur and laser light can be generated. To bring this about, opposite faces of the *p-n* junction crystal must be flat and parallel, so that light can be reflected back and forth within the crystal. (Recall that in the helium–neon laser of Fig. 40-20, a pair of mirrors served this purpose.) Thus, a *p-n* junction can act as a **junction laser**, its light output being highly coherent and much more sharply defined in wavelength than light from an LED.

Junction lasers are built into compact disc (CD) players, where, by detecting reflections from the rotating disc, they are used to translate microscopic pits in the disc into sound. They are also much used in optical communication systems based on optical fibers. Figure 41-18 suggests their tiny scale. Junction lasers are usually designed to operate in the infrared region of the electromagnetic spectrum because optical fibers have two “windows” in that region (at $\lambda = 1.31$ and $1.55 \mu\text{m}$) for which the energy absorption per unit length of the fiber is a minimum.



Fig. 41-18 A junction laser developed at the AT&T Bell Laboratories. The cube at the right is a grain of salt. (Courtesy AT&T Archives and History Center, Warren, NJ)

Sample Problem

Light-emitting diode (LED)

An LED is constructed from a *p-n* junction based on a certain Ga-As-P semiconducting material whose energy gap is 1.9 eV. What is the wavelength of the emitted light?

Calculation: For jumps from the bottom of the conduction band to the top of the valence band, Eq. 41-11 tells us

$$\lambda = \frac{hc}{E_g} = \frac{(6.63 \times 10^{-34} \text{ J}\cdot\text{s})(3.00 \times 10^8 \text{ m/s})}{(1.9 \text{ eV})(1.60 \times 10^{-19} \text{ J/eV})} \\ = 6.5 \times 10^{-7} \text{ m} = 650 \text{ nm.}$$

(Answer)

Light of this wavelength is red.



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41-11 The Transistor

A **transistor** is a three-terminal semiconducting device that can be used to amplify input signals. Figure 41-19 shows a generalized field-effect transistor (FET); in it, the flow of electrons from terminal *S* (the *source*) leftward through the shaded region to terminal *D* (the *drain*) can be controlled by an electric field (hence field effect) set up within the device by a suitable electric potential applied to terminal *G* (the *gate*). Transistors are available in many types; we shall discuss only a particular FET called a MOSFET, or metal-oxide-semiconductor-field-effect transistor. The MOSFET has been described as the workhorse of the modern electronics industry.

For many applications the MOSFET is operated in only two states: with the drain-to-source current I_{DS} ON (gate open) or with it OFF (gate closed). The first of these can represent a 1 and the other a 0 in the binary arithmetic on which digital logic is based, and therefore MOSFETs can be used in digital logic circuits. Switching between the ON and OFF states can occur at high speed, so that binary logic data can be moved through MOSFET-based circuits very rapidly. MOSFETs about 500 nm in length—about the same as the wavelength of yellow light—are routinely fabricated for use in electronic devices of all kinds.

Figure 41-20 shows the basic structure of a MOSFET. A single crystal of silicon or other semiconductor is lightly doped to form *p*-type material that serves as the *substrate*. Embedded in this substrate, by heavily “overdoping” with *n*-type dopants, are two “islands” of *n*-type material, forming the drain *D* and the source *S*. The drain and source are connected by a thin channel of *n*-type material, called the ***n* channel**. A thin insulating layer of silicon dioxide (hence the O in MOSFET) is deposited on the crystal and penetrated by two metallic terminals (hence the M) at *D* and *S*, so that electrical contact can be made with the drain and the source. A thin metallic layer—the gate *G*—is deposited facing the *n* channel.

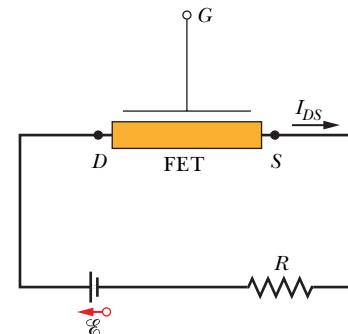


Fig. 41-19 A circuit containing a generalized field-effect transistor through which electrons flow from the source terminal *S* to the drain terminal *D*. (The conventional current I_{DS} is in the opposite direction.) The magnitude of I_{DS} is controlled by the electric field set up within the FET by a potential applied to *G*, the gate terminal.

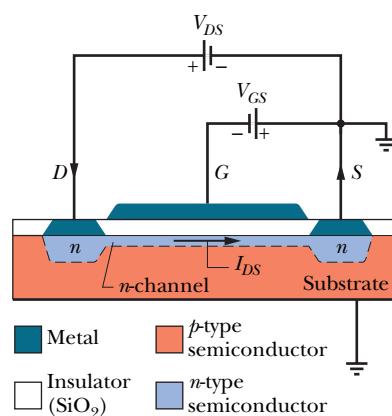


Fig. 41-20 A particular type of field-effect transistor known as a MOSFET. The magnitude of the drain-to-source conventional current I_{DS} through the *n* channel is controlled by the potential difference V_{GS} applied between the source *S* and the gate *G*. A depletion zone that exists between the *n*-type material and the *p*-type substrate is not shown.

Note that the gate makes no electrical contact with the transistor proper, being separated from it by the insulating oxide layer.

Consider first that the source and *p*-type substrate are grounded (at zero potential) and the gate is “floating”; that is, the gate is not connected to an external source of emf. Let a potential V_{DS} be applied between the drain and the source, such that the drain is positive. Electrons will then flow through the *n* channel from source to drain, and the conventional current I_{DS} , as shown in Fig. 41-20, will be from drain to source through the *n* channel.

Now let a potential V_{GS} be applied to the gate, making it negative with respect to the source. The negative gate sets up within the device an electric field (hence the “field effect”) that tends to repel electrons from the *n* channel down into the substrate. This electron movement widens the (naturally occurring) depletion zone between the *n* channel and the substrate, at the expense of the *n* channel. The reduced width of the *n* channel, coupled with a reduction in the number of charge carriers in that channel, increases the resistance of that channel and thus decreases the current I_{DS} . With the proper value of V_{GS} , this current can be shut off completely; hence, by controlling V_{GS} , the MOSFET can be switched between its ON and OFF modes.

Charge carriers do not flow through the substrate because it (1) is lightly doped, (2) is not a good conductor, and (3) is separated from the *n* channel and the two *n*-type islands by an insulating depletion zone, not specifically shown in Fig. 41-20. Such a depletion zone always exists at a boundary between *n*-type material and *p*-type material, as Fig. 41-12b shows.

Computers and other electronic devices employ thousands (if not millions) of transistors and other electronic components, such as capacitors and resistors. These are not assembled as separate units but are crafted into a single semiconducting **chip**, forming an **integrated circuit** with millions of transistors and many other electronic components.

REVIEW & SUMMARY

Metals, Semiconductors, and Insulators Three electrical properties that can be used to distinguish among crystalline solids are **resistivity** ρ , **temperature coefficient of resistivity** α , and **number density of charge carriers** n . Solids can be broadly divided into **insulators** (very high ρ), **metals** (low ρ , positive and low α , large n), and **semiconductors** (high ρ , negative and high α , small n).

Energy Levels and Gaps in a Crystalline Solid An isolated atom can exist in only a discrete set of energy levels. As atoms come together to form a solid, the levels of the individual atoms merge to form the discrete energy **bands** of the solid. These energy bands are separated by energy **gaps**, each of which corresponds to a range of energies that no electron may possess.

Any energy band is made up of an enormous number of very closely spaced levels. The Pauli exclusion principle asserts that only one electron may occupy each of these levels.

Insulators In an insulator, the highest band containing electrons is completely filled and is separated from the vacant band above it by an energy gap so large that electrons can essentially never become thermally agitated enough to jump across the gap.

Metals In a metal, the highest band that contains any electrons is only partially filled. The energy of the highest filled level at a temperature of 0 K is called the **Fermi energy** E_F for the metal; for copper, $E_F = 7.0 \text{ eV}$.

The electrons in the partially filled band are the **conduction electrons** and their number is

$$\left(\frac{\text{number of conduction electrons}}{\text{in sample}} \right) = \left(\frac{\text{number of atoms}}{\text{in sample}} \right) \times \left(\frac{\text{number of valence electrons per atom}}{\text{per atom}} \right). \quad (41-2)$$

The number of atoms in a sample is given by

$$\begin{aligned} \left(\frac{\text{number of atoms}}{\text{in sample}} \right) &= \frac{\text{sample mass } M_{\text{sam}}}{\text{atomic mass}} \\ &= \frac{\text{sample mass } M_{\text{sam}}}{(\text{molar mass } M)/N_A} \\ &= \frac{(\text{material's density})(\text{sample volume } V)}{(\text{molar mass } M)/N_A}. \end{aligned} \quad (41-4)$$

The number density n of the conduction electrons is

$$n = \frac{\text{number of conduction electrons in sample}}{\text{sample volume } V}. \quad (41-3)$$

The **density of states** function $N(E)$ is the number of available energy levels per unit volume of the sample and per unit energy in-

terval and is given by

$$N(E) = \frac{8\sqrt{2}\pi m^{3/2}}{h^3} E^{1/2} \quad (\text{density of states, } \text{m}^{-3} \text{ J}^{-1}), \quad (41-5)$$

where $m (= 9.109 \times 10^{-31} \text{ kg})$ is the electron mass, $h (= 6.626 \times 10^{-34} \text{ J} \cdot \text{s})$ is the Planck constant, and E is the energy in joules at which $N(E)$ is to be evaluated. To modify the equation so that the value of E is in eV and the value of $N(E)$ is in $\text{m}^{-3} \text{ eV}^{-1}$, multiply the right side by $e^{3/2}$ (where $e = 1.602 \times 10^{-19} \text{ C}$).

The **occupancy probability** $P(E)$, the probability that a given available state will be occupied by an electron, is

$$P(E) = \frac{1}{e^{(E-E_F)/kT} + 1} \quad (\text{occupancy probability}). \quad (41-6)$$

The **density of occupied states** $N_o(E)$ is given by the product of the two quantities in Eqs. (41-5) and (41-6):

$$N_o(E) = N(E) P(E) \quad (\text{density of occupied states}). \quad (41-7)$$

The Fermi energy for a metal can be found by integrating $N_o(E)$ for $T = 0$ from $E = 0$ to $E = E_F$. The result is

$$E_F = \left(\frac{3}{16\sqrt{2}\pi} \right)^{2/3} \frac{h^2}{m} n^{2/3} = \frac{0.121h^2}{m} n^{2/3}. \quad (41-9)$$

Semiconductors The band structure of a semiconductor is like that of an insulator except that the gap width E_g is much smaller in the semiconductor. For silicon (a semiconductor) at room temperature, thermal agitation raises a few electrons to the **conduction band**, leaving an equal number of **holes** in the **valence band**. Both electrons and holes serve as charge carriers.

The number of electrons in the conduction band of silicon can be increased greatly by doping with small amounts of phosphorus,

thus forming ***n*-type material**. The number of holes in the valence band can be greatly increased by doping with aluminum, thus forming ***p*-type material**.

The *p-n* Junction A ***p-n* junction** is a single semiconducting crystal with one end doped to form *p*-type material and the other end doped to form *n*-type material, the two types meeting at a **junction plane**. At thermal equilibrium, the following occurs at that plane:

The **majority carriers** (electrons on the *n* side and holes on the *p* side) diffuse across the junction plane, producing a **diffusion current** I_{diff} .

The **minority carriers** (holes on the *n* side and electrons on the *p* side) are swept across the junction plane, forming a **drift current** I_{drift} . These two currents are equal in magnitude, making the net current zero.

A **depletion zone**, consisting largely of charged donor and acceptor ions, forms across the junction plane.

A **contact potential difference** V_0 develops across the depletion zone.

Applications of the *p-n* Junction When a potential difference is applied across a *p-n* junction, the device conducts electricity more readily for one polarity of the applied potential difference than for the other. Thus, a *p-n* junction can serve as a **junction rectifier**.

When a *p-n* junction is forward biased, it can emit light, hence can serve as a **light-emitting diode** (LED). The wavelength of the emitted light is given by

$$\lambda = \frac{c}{f} = \frac{hc}{E_g}. \quad (41-11)$$

A strongly forward-biased *p-n* junction with parallel end faces can operate as a **junction laser**, emitting light of a sharply defined wavelength.

** View All Solutions Here **

Q U E S T I O N S

- 1** On which of the following does the interval between adjacent energy levels in the highest occupied band of a metal depend: (a) the material of which the sample is made, (b) the size of the sample, (c) the position of the level in the band, (d) the temperature of the sample, (e) the Fermi energy of the metal?
- 2** Figure 41-1a shows 14 atoms that represent the unit cell of copper. However, because each of these atoms is shared with one or more adjoining unit cells, only a fraction of each atom belongs to the unit cell shown. What is the number of atoms per unit cell for copper? (To answer, count up the fractional atoms belonging to a single unit cell.)
- 3** Figure 41-1b shows 18 atoms that represent the unit cell of silicon. Fourteen of these atoms, however, are shared with one or more adjoining unit cells. What is the number of atoms per unit cell for silicon? (See Question 2.)
- 4** Figure 41-21 shows three labeled levels in a band and also the Fermi level for the material. The temperature is 0 K. Rank the three levels according to the probability of occupation, greatest first if the temperature is (a) 0 K and (b) 1000 K.
-
- Fig. 41-21** Question 4.
- K. (c) At the latter temperature, rank the levels according to the density of states $N(E)$ there, greatest first.
- 5** The occupancy probability at a certain energy E_1 in the valence band of a metal is 0.60 when the temperature is 300 K. Is E_1 above or below the Fermi energy?
- 6** An isolated atom of germanium has 32 electrons, arranged in subshells according to this scheme:
- $$1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 3d^{10} \ 4s^2 \ 4p^2.$$
- This element has the same crystal structure as silicon and, like silicon, is a semiconductor. Which of these electrons form the valence band of crystalline germanium?
- 7** If the temperature of a piece of a metal is increased, does the probability of occupancy 0.1 eV above the Fermi level increase, decrease, or remain the same?
- 8** In the biased *p-n* junctions shown in Fig. 41-15, there is an electric field \vec{E} in each of the two depletion zones, associated with the potential difference that exists across that zone. (a) Is the electric field vector directed from left to right in the figure or from right to left? (b) Is the magnitude of the field greater for forward bias or for back bias?

** View All Solutions Here **

1162 CHAPTER 41 CONDUCTION OF ELECTRICITY IN SOLIDS

9 Consider a copper wire that is carrying, say, a few amperes of current. Is the drift speed v_d of the conduction electrons that form that current about equal to, much greater than, or much less than the Fermi speed v_F for copper (the speed associated with the Fermi energy for copper)?

10 In a silicon lattice, where should you look if you want to find (a) a conduction electron, (b) a valence electron, and (c) an electron associated with the $2p$ subshell of the isolated silicon atom?

11 The energy gaps E_g for the semiconductors silicon and germanium are, respectively, 1.12 and 0.67 eV. Which of the following statements, if any, are true? (a) Both substances have the same number density of charge carriers at room temperature. (b) At room temperature, germanium has a greater number density of charge carriers than silicon. (c) Both substances have a greater number density of conduction electrons than holes. (d) For each substance, the number density of electrons equals that of holes.

PROBLEMS



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>

sec. 41-5 Metals

•1 Show that Eq. 41-9 can be written as $E_F = An^{2/3}$, where the constant A has the value $3.65 \times 10^{-19} \text{ m}^2 \cdot \text{eV}$.

•2 Calculate the density of states $N(E)$ for a metal at energy $E = 8.0 \text{ eV}$ and show that your result is consistent with the curve of Fig. 41-6.

•3 Copper, a monovalent metal, has molar mass 63.54 g/mol and density 8.96 g/cm^3 . What is the number density n of conduction electrons in copper?

•4 A state 63 meV above the Fermi level has a probability of occupancy of 0.090 . What is the probability of occupancy for a state 63 meV below the Fermi level?

•5 (a) Show that Eq. 41-5 can be written as $N(E) = CE^{1/2}$. (b) Evaluate C in terms of meters and electron-volts. (c) Calculate $N(E)$ for $E = 5.00 \text{ eV}$.

•6 Use Eq. 41-9 to verify 7.0 eV as copper's Fermi energy.

•7 SSM What is the probability that a state 0.0620 eV above the Fermi energy will be occupied at (a) $T = 0 \text{ K}$ and (b) $T = 320 \text{ K}$?

•8 What is the number density of conduction electrons in gold, which is a monovalent metal? Use the molar mass and density provided in Appendix F.

•9 SSM WWW Silver is a monovalent metal. Calculate (a) the number density of conduction electrons, (b) the Fermi energy, (c) the Fermi speed, and (d) the de Broglie wavelength corresponding to this electron speed. See Appendix F for the needed data on silver.

•10 Show that the probability $P(E)$ that an energy level having energy E is not occupied is

$$P(E) = \frac{1}{e^{-\Delta E/kT} + 1},$$

where $\Delta E = E - E_F$.

•11 Calculate $N_o(E)$, the density of occupied states, for copper at $T = 1000 \text{ K}$ for an energy E of (a) 4.00 eV , (b) 6.75 eV , (c) 7.00 eV , (d) 7.25 eV , and (e) 9.00 eV . Compare your results with the graph of Fig. 41-8b. The Fermi energy for copper is 7.00 eV .

•12 What is the probability that, at a temperature of $T = 300 \text{ K}$, an electron will jump across the energy gap $E_g (= 5.5 \text{ eV})$ in a dia-

mond that has a mass equal to the mass of Earth? Use the molar mass of carbon in Appendix F; assume that in diamond there is one valence electron per carbon atom.

•13 GO The Fermi energy for copper is 7.00 eV . For copper at 1000 K , (a) find the energy of the energy level whose probability of being occupied by an electron is 0.900 . For this energy, evaluate (b) the density of states $N(E)$ and (c) the density of occupied states $N_o(E)$.

•14 Assume that the total volume of a metal sample is the sum of the volume occupied by the metal ions making up the lattice and the (separate) volume occupied by the conduction electrons. The density and molar mass of sodium (a metal) are 971 kg/m^3 and 23.0 g/mol , respectively; assume the radius of the Na^+ ion is 98.0 pm . (a) What percent of the volume of a sample of metallic sodium is occupied by its conduction electrons? (b) Carry out the same calculation for copper, which has density, molar mass, and ionic radius of 8960 kg/m^3 , 63.5 g/mol , and 135 pm , respectively. (c) For which of these metals do you think the conduction electrons behave more like a free-electron gas?

•15 SSM WWW In Eq. 41-6 let $E - E_F = \Delta E = 1.00 \text{ eV}$. (a) At what temperature does the result of using this equation differ by 1.0% from the result of using the classical Boltzmann equation $P(E) = e^{-\Delta E/kT}$ (which is Eq. 41-1 with two changes in notation)? (b) At what temperature do the results from these two equations differ by 10% ?

•16 Calculate the number density (number per unit volume) for (a) molecules of oxygen gas at 0.0°C and 1.0 atm pressure and (b) conduction electrons in copper. (c) What is the ratio of the latter to the former? What is the average distance between (d) the oxygen molecules and (e) the conduction electrons, assuming this distance is the edge length of a cube with a volume equal to the available volume per particle (molecule or electron)?

•17 The Fermi energy of aluminum is 11.6 eV ; its density and molar mass are 2.70 g/cm^3 and 27.0 g/mol , respectively. From these data, determine the number of conduction electrons per atom.

•18 GO A sample of a certain metal has a volume of $4.0 \times 10^{-5} \text{ m}^3$. The metal has a density of 9.0 g/cm^3 and a molar mass of 60 g/mol . The atoms are bivalent. How many conduction electrons (or valence electrons) are in the sample?

•19 The Fermi energy for silver is 5.5 eV . At $T = 0^\circ\text{C}$, what are

the probabilities that states with the following energies are occupied: (a) 4.4 eV, (b) 5.4 eV, (c) 5.5 eV, (d) 5.6 eV, and (e) 6.4 eV? (f) At what temperature is the probability 0.16 that a state with energy $E = 5.6$ eV is occupied?

••20 GO What is the number of occupied states in the energy range of 0.0300 eV that is centered at a height of 6.10 eV in the valence band if the sample volume is 5.00×10^{-8} m³, the Fermi level is 5.00 eV, and the temperature is 1500 K?

••21 At 1000 K, the fraction of the conduction electrons in a metal that have energies greater than the Fermi energy is equal to the area under the curve of Fig. 41-8b beyond E_F divided by the area under the entire curve. It is difficult to find these areas by direct integration. However, an approximation to this fraction at any temperature T is

$$\text{frac} = \frac{3kT}{2E_F}.$$

Note that $\text{frac} = 0$ for $T = 0$ K, just as we would expect. What is this fraction for copper at (a) 300 K and (b) 1000 K? For copper, $E_F = 7.0$ eV. (c) Check your answers by numerical integration using Eq. 41-7.

••22 At what temperature do 1.30% of the conduction electrons in lithium (a metal) have energies greater than the Fermi energy E_F , which is 4.70 eV? (See Problem 21.)

••23 Show that, at $T = 0$ K, the average energy E_{avg} of the conduction electrons in a metal is equal to $\frac{2}{3}E_F$. (*Hint:* By definition of average, $E_{\text{avg}} = (1/n) \int E N_o(E) dE$, where n is the number density of charge carriers.)

••24 GO A certain material has a molar mass of 20.0 g/mol, a Fermi energy of 5.00 eV, and 2 valence electrons per atom. What is the density (g/cm³)?

••25 (a) Using the result of Problem 23 and 7.00 eV for copper's Fermi energy, determine how much energy would be released by the conduction electrons in a copper coin with mass 3.10 g if we could suddenly turn off the Pauli exclusion principle. (b) For how long would this amount of energy light a 100 W lamp? (*Note:* There is no way to turn off the Pauli principle!)

••26 At $T = 300$ K, how far above the Fermi energy is a state for which the probability of occupation by a conduction electron is 0.10?

••27 Zinc is a bivalent metal. Calculate (a) the number density of conduction electrons, (b) the Fermi energy, (c) the Fermi speed, and (d) the de Broglie wavelength corresponding to this electron speed. See Appendix F for the needed data on zinc.

••28 GO What is the Fermi energy of gold (a monovalent metal with molar mass 197 g/mol and density 19.3 g/cm³)?

••29 Use the result of Problem 23 to calculate the total translational kinetic energy of the conduction electrons in 1.00 cm³ of copper at $T = 0$ K.

••30 GO A certain metal has 1.70×10^{28} conduction electrons per cubic meter. A sample of that metal has a volume of 6.00×10^{-6} m³ and a temperature of 200 K. How many occupied states are in the energy range of 3.20×10^{-20} J that is centered on the energy 4.00×10^{-19} J? (*Caution:* Avoid round-off in the exponential.)

sec. 41-6 Semiconductors

•31 SSM (a) What maximum light wavelength will excite an electron in the valence band of diamond to the conduction band?

The energy gap is 5.50 eV. (b) In what part of the electromagnetic spectrum does this wavelength lie?

••32 The compound gallium arsenide is a commonly used semiconductor, having an energy gap E_g of 1.43 eV. Its crystal structure is like that of silicon, except that half the silicon atoms are replaced by gallium atoms and half by arsenic atoms. Draw a flattened-out sketch of the gallium arsenide lattice, following the pattern of Fig. 41-10a. What is the net charge of the (a) gallium and (b) arsenic ion core? (c) How many electrons per bond are there? (*Hint:* Consult the periodic table in Appendix G.)

••33 The occupancy probability function (Eq. 41-6) can be applied to semiconductors as well as to metals. In semiconductors the Fermi energy is close to the midpoint of the gap between the valence band and the conduction band. For germanium, the gap width is 0.67 eV. What is the probability that (a) a state at the bottom of the conduction band is occupied and (b) a state at the top of the valence band is not occupied? Assume that $T = 290$ K. (*Note:* In a pure semiconductor, the Fermi energy lies symmetrically between the population of conduction electrons and the population of holes and thus is at the center of the gap. There need not be an available state at the location of the Fermi energy.)

••34 In a simplified model of an undoped semiconductor, the actual distribution of energy states may be replaced by one in which there are N_v states in the valence band, all these states having the same energy E_v , and N_c states in the conduction band, all these states having the same energy E_c . The number of electrons in the conduction band equals the number of holes in the valence band. (a) Show that this last condition implies that

$$\frac{N_c}{\exp(\Delta E_c/kT) + 1} = \frac{N_v}{\exp(\Delta E_v/kT) + 1},$$

in which

$$\Delta E_c = E_c - E_F \quad \text{and} \quad \Delta E_v = -(E_v - E_F).$$

(b) If the Fermi level is in the gap between the two bands and its distance from each band is large relative to kT , then the exponentials dominate in the denominators. Under these conditions, show that

$$E_F = \frac{(E_c + E_v)}{2} + \frac{kT \ln(N_v/N_c)}{2}$$

and that, if $N_v \approx N_c$, the Fermi level for the undoped semiconductor is close to the gap's center.

sec. 41-7 Doped Semiconductors

••35 SSM WWW What mass of phosphorus is needed to dope 1.0 g of silicon so that the number density of conduction electrons in the silicon is increased by a multiply factor of 10^6 from the 10^{16} m⁻³ in pure silicon.

••36 A silicon sample is doped with atoms having donor states 0.110 eV below the bottom of the conduction band. (The energy gap in silicon is 1.11 eV.) If each of these donor states is occupied with a probability of 5.00×10^{-5} at $T = 300$ K, (a) is the Fermi level above or below the top of the silicon valence band and (b) how far above or below? (c) What then is the probability that a state at the bottom of the silicon conduction band is occupied?

••37 Doping changes the Fermi energy of a semiconductor. Consider silicon, with a gap of 1.11 eV between the top of the va-

1164 CHAPTER 41 CONDUCTION OF ELECTRICITY IN SOLIDS

lence band and the bottom of the conduction band. At 300 K the Fermi level of the pure material is nearly at the mid-point of the gap. Suppose that silicon is doped with donor atoms, each of which has a state 0.15 eV below the bottom of the silicon conduction band, and suppose further that doping raises the Fermi level to 0.11 eV below the bottom of that band (Fig. 41-22). For (a) pure and (b) doped silicon, calculate the probability that a state at the bottom of the silicon conduction band is occupied. (c) Calculate the probability that a state in the doped material (at the donor level) is occupied.

•38 Pure silicon at room temperature has an electron number density in the conduction band of about $5 \times 10^{15} \text{ m}^{-3}$ and an equal density of holes in the valence band. Suppose that one of every 10^7 silicon atoms is replaced by a phosphorus atom. (a) Which type will the doped semiconductor be, n or p ? (b) What charge carrier number density will the phosphorus add? (c) What is the ratio of the charge carrier number density (electrons in the conduction band and holes in the valence band) in the doped silicon to that in pure silicon?

sec. 41-9 The Junction Rectifier

•39 SSM When a photon enters the depletion zone of a p - n junction, the photon can scatter from the valence electrons there, transferring part of its energy to each electron, which then jumps to the conduction band. Thus, the photon creates electron–hole pairs. For this reason, the junctions are often used as light detectors, especially in the x-ray and gamma-ray regions of the electromagnetic spectrum. Suppose a single 662 keV gamma-ray photon transfers its energy to electrons in multiple scattering events inside a semiconductor with an energy gap of 1.1 eV, until all the energy is transferred. Assuming that each electron jumps the gap from the top of the valence band to the bottom of the conduction band, find the number of electron–hole pairs created by the process.

•40 For an ideal p - n junction rectifier with a sharp boundary between its two semiconducting sides, the current I is related to the potential difference V across the rectifier by

$$I = I_0(e^{eV/kT} - 1),$$

where I_0 , which depends on the materials but not on I or V , is called the *reverse saturation current*. The potential difference V is positive if the rectifier is forward-biased and negative if it is back-biased. (a) Verify that this expression predicts the behavior of a junction rectifier by graphing I versus V from -0.12 V to $+0.12 \text{ V}$. Take $T = 300 \text{ K}$ and $I_0 = 5.0 \text{ nA}$. (b) For the same temperature, calculate the ratio of the current for a 0.50 V forward bias to the current for a 0.50 V back bias.

sec. 41-10 The Light-Emitting Diode (LED)

•41 In a particular crystal, the highest occupied band is full. The crystal is transparent to light of wavelengths longer than 295 nm but opaque at shorter wavelengths. Calculate, in electron-volts, the gap between the highest occupied band and the next higher (empty) band for this material.

•42 A potassium chloride crystal has an energy band gap of 7.6

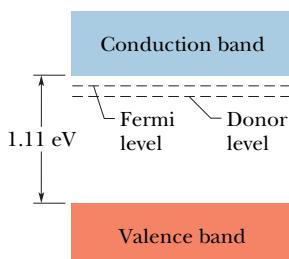


Fig. 41-22 Problem 37.

eV above the topmost occupied band, which is full. Is this crystal opaque or transparent to light of wavelength 140 nm ?

sec. 41-11 The Transistor

•43 A certain computer chip that is about the size of a postage stamp ($2.54 \text{ cm} \times 2.22 \text{ cm}$) contains about 3.5 million transistors. If the transistors are square, what must be their *maximum* dimension? (Note: Devices other than transistors are also on the chip, and there must be room for the interconnections among the circuit elements. Transistors smaller than $0.7 \mu\text{m}$ are now commonly and inexpensively fabricated.)

•44 A silicon-based MOSFET has a square gate $0.50 \mu\text{m}$ on edge. The insulating silicon oxide layer that separates the gate from the p -type substrate is $0.20 \mu\text{m}$ thick and has a dielectric constant of 4.5. (a) What is the equivalent gate–substrate capacitance (treating the gate as one plate and the substrate as the other plate)? (b) Approximately how many elementary charges e appear in the gate when there is a gate–source potential difference of 1.0 V ?

Additional Problems

45 SSM (a) Show that the slope dP/dE of Eq. 41-6 evaluated at $E = E_F$ is $-1/4kT$. (b) Show that the tangent line to the curve of Fig. 41-7b evaluated at $E = E_F$ intercepts the horizontal axis at $E = E_F + 2kT$.

46 Calculate dp/dT at room temperature for (a) copper and (b) silicon, using data from Table 41-1.

47 (a) Find the angle θ between adjacent nearest-neighbor bonds in the silicon lattice. Recall that each silicon atom is bonded to four of its nearest neighbors. The four neighbors form a regular tetrahedron—a pyramid whose sides and base are equilateral triangles. (b) Find the bond length, given that the atoms at the corners of the tetrahedron are 388 pm apart.

48 Show that $P(E)$, the occupancy probability in Eq. 41-6, is symmetrical about the value of the Fermi energy; that is, show that

$$P(E_F + \Delta E) + P(E_F - \Delta E) = 1.$$

49 (a) Show that the density of states at the Fermi energy is given by

$$N(E_F) = \frac{(4)(3^{1/3})(\pi^{2/3})mn^{1/3}}{h^2} \\ = (4.11 \times 10^{18} \text{ m}^{-2} \text{ eV}^{-1})n^{1/3},$$

in which n is the number density of conduction electrons. (b) Calculate $N(E_F)$ for copper, which is a monovalent metal with molar mass 63.54 g/mol and density 8.96 g/cm^3 . (c) Verify your calculation with the curve of Fig. 41-6, recalling that $E_F = 7.0 \text{ eV}$ for copper.

50 Silver melts at 961°C . At the melting point, what fraction of the conduction electrons are in states with energies greater than the Fermi energy of 5.5 eV ? (See Problem 21.)

51 The Fermi energy of copper is 7.0 eV . Verify that the corresponding Fermi speed is 1600 km/s .

52 Verify the numerical factor 0.121 in Eq. 41-9.

53 At what pressure, in atmospheres, would the number of molecules per unit volume in an ideal gas be equal to the number density of the conduction electrons in copper, with both gas and copper at temperature $T = 300 \text{ K}$?

NUCLEAR PHYSICS

42

42-1 WHAT IS PHYSICS?

Thus far in our exploration of the quantum amusement park, we have examined electrons trapped within various potential wells, including atoms, but we have largely neglected what lies at the center of an atom—the nucleus. For the last 90 years, a principal goal of physics has been to work out the quantum physics of nuclei, and, for almost as long, a principal goal of some types of engineering has been to apply that quantum physics with applications ranging from radiation therapy in the war on cancer to detectors of radon gas in basements.

Before we get to such applications and the quantum physics of nuclei, let's first discuss how physicists discovered that an atom has a nucleus. As obvious as that fact is today, it initially came as an incredible surprise.

42-2 Discovering the Nucleus

In the first years of the 20th century, not much was known about the structure of atoms beyond the fact that they contain electrons. The electron had been discovered (by J. J. Thomson) in 1897, and its mass was unknown in those early days. Thus, it was not possible even to say how many negatively charged electrons a given atom contained. Scientists reasoned that because atoms were electrically neutral, they must also contain some positive charge, but nobody knew what form this compensating positive charge took. One popular model was that the positive and negative charges were spread uniformly in a sphere.

In 1911 Ernest Rutherford proposed that the positive charge of the atom is densely concentrated at the center of the atom, forming its **nucleus**, and that, furthermore, the nucleus is responsible for most of the mass of the atom. Rutherford's proposal was no mere conjecture but was based firmly on the results of an experiment suggested by him and carried out by his collaborators, Hans Geiger (of Geiger counter fame) and Ernest Marsden, a 20-year-old student who had not yet earned his bachelor's degree.

In Rutherford's day it was known that certain elements, called **radioactive**, transform into other elements spontaneously, emitting particles in the process. One such element is radon, which emits alpha (α) particles that have an energy of about 5.5 MeV. We now know that these particles are helium nuclei.

Rutherford's idea was to direct energetic alpha particles at a thin target foil and measure the extent to which they were deflected as they passed through the foil. Alpha particles, which are about 7300 times more massive than electrons, have a charge of $+2e$.

Figure 42-1 shows the experimental arrangement of Geiger and Marsden. Their alpha source was a thin-walled glass tube of radon gas. The experiment involves counting the number of alpha particles that are deflected through various scattering angles ϕ .

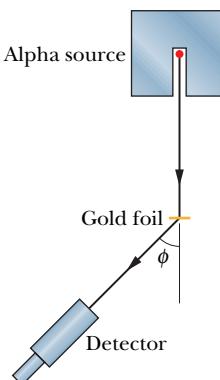


Fig. 42-1 An arrangement (top view) used in Rutherford's laboratory in 1911–1913 to study the scattering of α particles by thin metal foils. The detector can be rotated to various values of the scattering angle ϕ . The alpha source was radon gas, a decay product of radium. With this simple “tabletop” apparatus, the atomic nucleus was discovered.

Fig. 42-2 The dots are alpha-particle scattering data for a gold foil, obtained by Geiger and Marsden using the apparatus of Fig. 42-1. The solid curve is the theoretical prediction, based on the assumption that the atom has a small, massive, positively charged nucleus. The data have been adjusted to fit the theoretical curve at the experimental point that is enclosed in a circle.

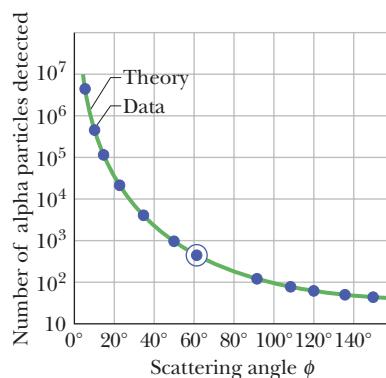


Figure 42-2 shows their results. Note especially that the vertical scale is logarithmic. We see that most of the particles are scattered through rather small angles, but—and this was the big surprise—a very small fraction of them are scattered through very large angles, approaching 180° . In Rutherford's words: "It was quite the most incredible event that ever happened to me in my life. It was almost as incredible as if you had fired a 15-inch shell at a piece of tissue paper and it [the shell] came back and hit you."

Why was Rutherford so surprised? At the time of these experiments, most physicists believed in the so-called plum pudding model of the atom, which had been advanced by J. J. Thomson. In this view the positive charge of the atom was thought to be spread out through the entire volume of the atom. The electrons (the "plums") were thought to vibrate about fixed points within this sphere of positive charge (the "pudding").

The maximum deflecting force that could act on an alpha particle as it passed through such a large positive sphere of charge would be far too small to deflect the alpha particle by even as much as 1° . (The expected deflection has been compared to what you would observe if you fired a bullet through a sack of snowballs.) The electrons in the atom would also have very little effect on the massive, energetic alpha particle. They would, in fact, be themselves strongly deflected, much as a swarm of gnats would be brushed aside by a stone thrown through them.

Rutherford saw that, to deflect the alpha particle backward, there must be a large force; this force could be provided if the positive charge, instead of being spread throughout the atom, were concentrated tightly at its center. Then the incoming alpha particle could get very close to the positive charge without penetrating it; such a close encounter would result in a large deflecting force.

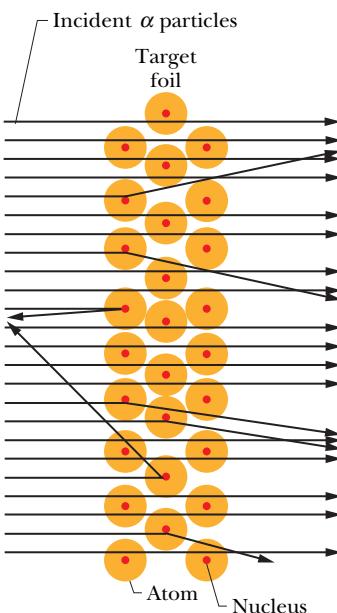


Fig. 42-3 The angle through which an incident alpha particle is scattered depends on how close the particle's path lies to an atomic nucleus. Large deflections result only from very close encounters.

Figure 42-3 shows possible paths taken by typical alpha particles as they pass through the atoms of the target foil. As we see, most are either undeflected or only slightly deflected, but a few (those whose incoming paths pass, by chance, very close to a nucleus) are deflected through large angles. From an analysis of the data, Rutherford concluded that the radius of the nucleus must be smaller than the radius of an atom by a factor of about 10^4 . In other words, the atom is mostly empty space.

Sample Problem**Rutherford scattering of an alpha particle by a gold nucleus**

An alpha particle with kinetic energy $K_i = 5.30 \text{ MeV}$ happens, by chance, to be headed directly toward the nucleus of a neutral gold atom (Fig. 42-4a). What is its *distance of closest approach d* (least center-to-center separation) to the nucleus? Assume that the atom remains stationary.

KEY IDEAS

(1) Throughout the motion, the total mechanical energy E of the particle–atom system is conserved. (2) In addition to the kinetic energy, that total energy includes electric potential energy U as given by Eq. 24-43 ($U = q_1 q_2 / 4\pi\epsilon_0 r$).

Calculations: The alpha particle has a charge of $+2e$ because it contains two protons. The target nucleus has a charge of $q_{\text{Au}} = +79e$ because it contains 79 protons. However, that nuclear charge is surrounded by an electron “cloud” with a charge of $q_e = -79e$, and thus the alpha particle initially “sees” a neutral atom with a net charge of $q_{\text{atom}} = 0$. The electric force on the particle and the initial electric potential energy of the particle–atom system is $U_i = 0$.

Once the alpha particle enters the atom, we say that it passes through the electron cloud surrounding the nucleus.

That cloud then acts as a closed conducting spherical shell and, by Gauss’ law, has no effect on the (now internal) charged alpha particle. Then the alpha particle “sees” only the nuclear charge q_{Au} . Because q_α and q_{Au} are both positively charged, a repulsive electric force acts on the alpha particle, slowing it, and the particle–atom system has a potential energy

$$U = \frac{1}{4\pi\epsilon_0} \frac{q_\alpha q_{\text{Au}}}{r}$$

that depends on the center-to-center separation r of the incoming particle and the target nucleus (Fig. 42-4b).

As the repulsive force slows the alpha particle, energy is transferred from kinetic energy to electric potential energy. The transfer is complete when the alpha particle momentarily stops at the distance of closest approach d to the target nucleus (Fig. 42-4c). Just then the kinetic energy is $K_f = 0$ and the particle–atom system has the electric potential energy

$$U_f = \frac{1}{4\pi\epsilon_0} \frac{q_\alpha q_{\text{Au}}}{d}$$

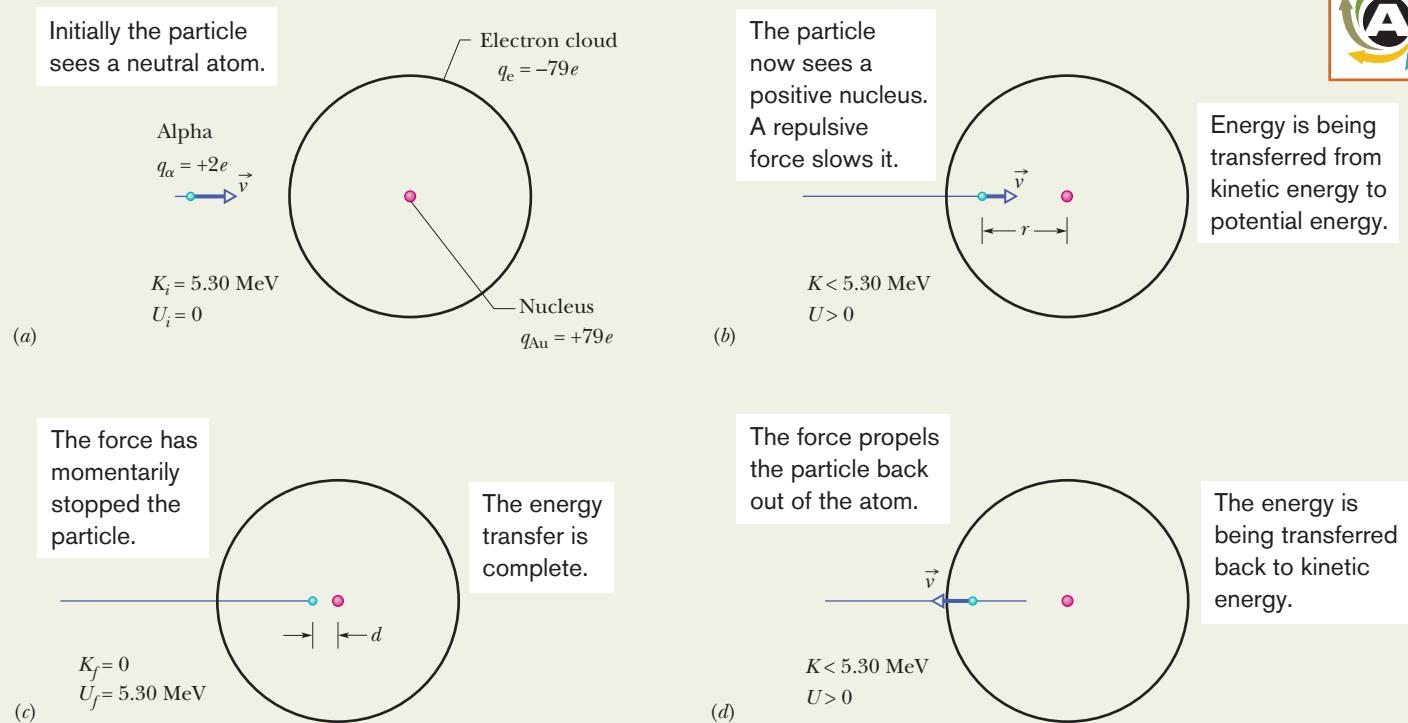


Fig. 42-4 An alpha particle (a) approaches and (b) then enters a gold atom, headed toward the nucleus. The alpha particle (c) comes to a stop at the point of closest approach and (d) is propelled back out of the atom.

To find d , we conserve the total mechanical energy between the initial state i and this later state f , writing

$$K_i + U_i = K_f + U_f$$

and $K_i + 0 = 0 + \frac{1}{4\pi\epsilon_0} \frac{q_\alpha q_{\text{Au}}}{d}$.

(We are assuming that the alpha particle is not affected by the force holding the nucleus together, which acts over only a short distance.) Solving for d and then substituting for the charges and initial kinetic energy lead to

$$\begin{aligned} d &= \frac{(2e)(79e)}{4\pi\epsilon_0 K_\alpha} \\ &= \frac{(2 \times 79)(1.60 \times 10^{-19} \text{ C})^2}{4\pi\epsilon_0 (5.30 \text{ MeV})(1.60 \times 10^{-13} \text{ J/MeV})} \\ &= 4.29 \times 10^{-14} \text{ m.} \end{aligned} \quad (\text{Answer})$$

This distance is considerably larger than the sum of the radii of the gold nucleus and the alpha particle. Thus, this alpha particle reverses its motion (Fig. 42-4d) without ever actually “touching” the gold nucleus.



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42-3 Some Nuclear Properties

Table 42-1 shows some properties of a few atomic nuclei. When we are interested primarily in their properties as specific nuclear species (rather than as parts of atoms), we call these particles **nuclides**.

Some Nuclear Terminology

Nuclei are made up of protons and neutrons. The number of protons in a nucleus (called the **atomic number** or **proton number** of the nucleus) is represented by the symbol Z ; the number of neutrons (the **neutron number**) is represented by the symbol N . The total number of neutrons and protons in a nucleus is called its **mass number** A ; thus

$$A = Z + N. \quad (42-1)$$

Neutrons and protons, when considered collectively, are called **nucleons**.

We represent nuclides with symbols such as those displayed in the first column of Table 42-1. Consider ^{197}Au , for example. The superscript 197 is the mass

Table 42-1

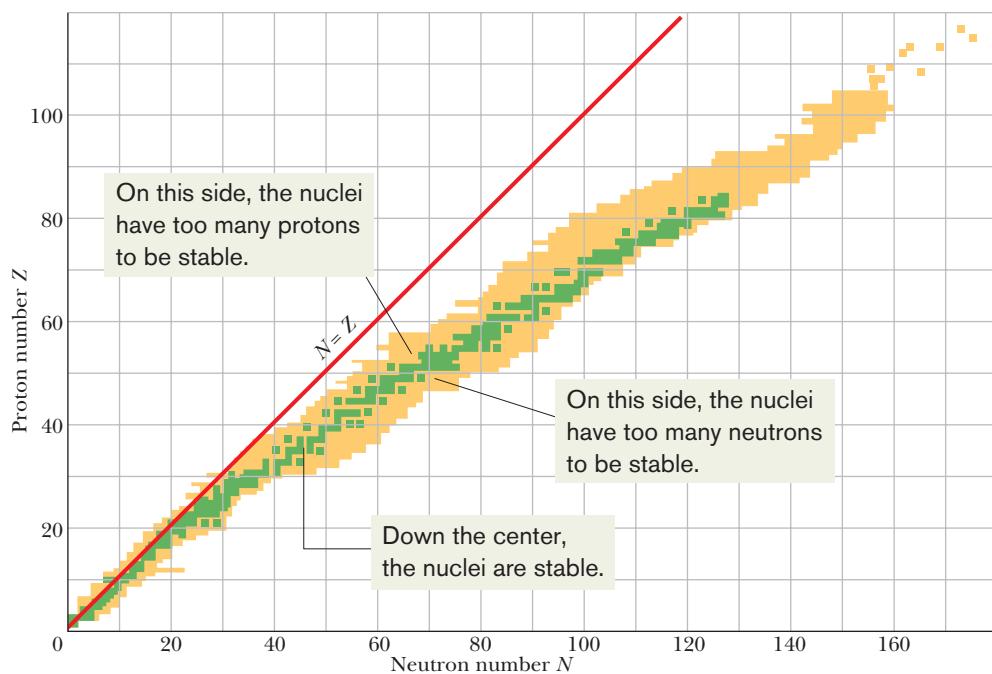
Some Properties of Selected Nuclides

Nuclide	Z	N	A	Stability ^a	Mass ^b (u)	Spin ^c	Binding Energy (MeV/nucleon)
^1H	1	0	1	99.985%	1.007 825	$\frac{1}{2}$	—
^7Li	3	4	7	92.5%	7.016 004	$\frac{3}{2}$	5.60
^{31}P	15	16	31	100%	30.973 762	$\frac{1}{2}$	8.48
^{84}Kr	36	48	84	57.0%	83.911 507	0	8.72
^{120}Sn	50	70	120	32.4%	119.902 197	0	8.51
^{157}Gd	64	93	157	15.7%	156.923 957	$\frac{3}{2}$	8.21
^{197}Au	79	118	197	100%	196.966 552	$\frac{3}{2}$	7.91
^{227}Ac	89	138	227	21.8 y	227.027 747	$\frac{3}{2}$	7.65
^{239}Pu	94	145	239	24 100 y	239.052 157	$\frac{1}{2}$	7.56

^aFor stable nuclides, the **isotopic abundance** is given; this is the fraction of atoms of this type found in a typical sample of the element. For radioactive nuclides, the half-life is given.

^bFollowing standard practice, the reported mass is that of the neutral atom, not that of the bare nucleus.

^cSpin angular momentum in units of \hbar .



number A . The chemical symbol Au tells us that this element is gold, whose atomic number is 79. From Eq. 42-1, the neutron number of this nuclide is $197 - 79$, or 118.

Nuclides with the same atomic number Z but different neutron numbers N are called **isotopes** of one another. The element gold has 32 isotopes, ranging from ^{173}Au to ^{204}Au . Only one of them (^{197}Au) is stable; the remaining 31 are radioactive. Such **radionuclides** undergo **decay** (or **disintegration**) by emitting a particle and thereby transforming to a different nuclide.

Organizing the Nuclides

The neutral atoms of all isotopes of an element (all with the same Z) have the same number of electrons and the same chemical properties, and they fit into the same box in the periodic table of the elements. The *nuclear* properties of the isotopes of a given element, however, are very different from one isotope to another. Thus, the periodic table is of limited use to the nuclear physicist, the nuclear chemist, or the nuclear engineer.

We organize the nuclides on a **nuclidic chart** like that in Fig. 42-5, in which a nuclide is represented by plotting its proton number against its neutron number. The stable nuclides in this figure are represented by the green, the radionuclides by the beige. As you can see, the radionuclides tend to lie on either side of—and at the upper end of—a well-defined band of stable nuclides. Note too that light stable nuclides tend to lie close to the line $N = Z$, which means that they have about the same numbers of neutrons and protons. Heavier nuclides, however, tend to have many more neutrons than protons. As an example, we saw that ^{197}Au has 118 neutrons and only 79 protons, a *neutron excess* of 39.

Nuclidic charts are available as wall charts, in which each small box on the chart is filled with data about the nuclide it represents. Figure 42-6 shows a section of such a chart, centered on ^{197}Au .

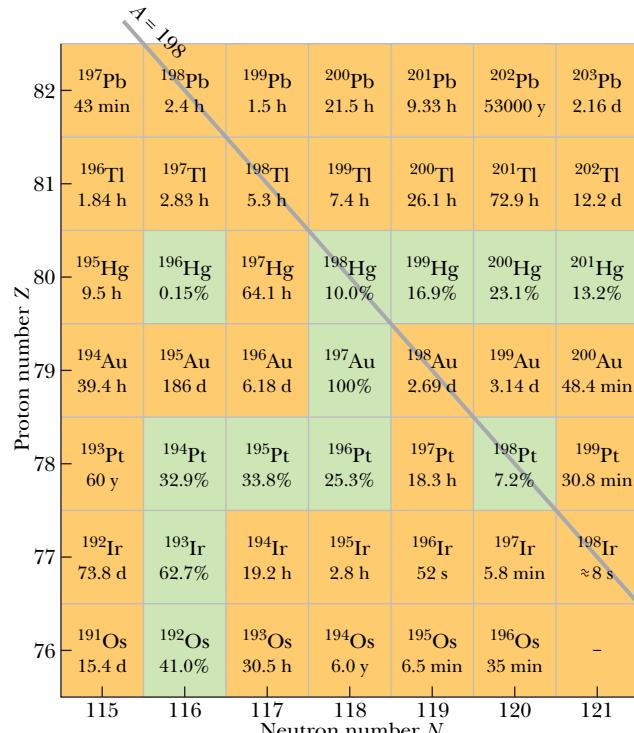


Fig. 42-6 An enlarged and detailed section of the nuclidic chart of Fig. 42-5, centered on ^{197}Au . Green squares represent stable nuclides, for which relative isotopic abundances are given. Beige squares represent radionuclides, for which half-lives are given. Isobaric lines of constant mass number A slope as shown by the example line for $A = 198$.

Relative abundances (usually, as found on Earth) are shown for stable nuclides, and half-lives (a measure of decay rate) are shown for radionuclides. The sloping line points out a line of **isobars**—nuclides of the same mass number, $A = 198$ in this case.

In recent years, nuclides with atomic numbers as high as $Z = 118$ ($A = 294$) have been found in laboratory experiments (no elements with Z greater than 92 occur naturally). Although large nuclides generally should be highly unstable and last only a very brief time, certain supermassive nuclides are relatively stable, with fairly long lifetimes. These stable supermassive nuclides and other predicted ones form an *island of stability* at high values of Z and N on a nuclidic chart like Fig. 42-5.

CHECKPOINT 1

Based on Fig. 42-5, which of the following nuclides do you conclude are not likely to be detected: ^{52}Fe ($Z = 26$), ^{90}As ($Z = 33$), ^{158}Nd ($Z = 60$), ^{175}Lu ($Z = 71$), ^{208}Pb ($Z = 82$)?

Nuclear Radii

A convenient unit for measuring distances on the scale of nuclei is the *femtometer* ($= 10^{-15} \text{ m}$). This unit is often called the *fermi*; the two names share the same abbreviation. Thus,

$$1 \text{ femtometer} = 1 \text{ fermi} = 1 \text{ fm} = 10^{-15} \text{ m}. \quad (42-2)$$

We can learn about the size and structure of nuclei by bombarding them with a beam of high-energy electrons and observing how the nuclei deflect the incident electrons. The electrons must be energetic enough (at least 200 MeV) to have de Broglie wavelengths that are smaller than the nuclear structures they are to probe.

The nucleus, like the atom, is not a solid object with a well-defined surface. Furthermore, although most nuclides are spherical, some are notably ellipsoidal. Nevertheless, electron-scattering experiments (as well as experiments of other kinds) allow us to assign to each nuclide an effective radius given by

$$r = r_0 A^{1/3}, \quad (42-3)$$

in which A is the mass number and $r_0 \approx 1.2 \text{ fm}$. We see that the volume of a nucleus, which is proportional to r^3 , is directly proportional to the mass number A and is independent of the separate values of Z and N . That is, we can treat most nuclei as being a sphere with a volume that depends on the number of nucleons, regardless of their type.

Equation 42-3 does not apply to *halo nuclides*, which are neutron-rich nuclides that were first produced in laboratories in the 1980s. These nuclides are larger than predicted by Eq. 42-3, because some of the neutrons form a *halo* around a spherical core of the protons and the rest of the neutrons. Lithium isotopes give an example. When a neutron is added to ^6Li to form ^7Li , neither of which are halo nuclides, the effective radius increases by about 4%. However, when two neutrons are added to ^6Li to form the neutron-rich isotope ^{11}Li (the largest of the lithium isotopes), they do not join that existing nucleus but instead form a halo around it, increasing the effective radius by about 30%. Apparently this halo configuration involves less energy than a core containing all 11 nucleons. (In this chapter we shall generally assume that Eq. 42-3 applies.)

Atomic Masses

Atomic masses are now measured to great precision, but usually nuclear masses are not directly measurable because stripping off all the electrons from an atom is difficult. As we briefly discussed in Section 37-12, atomic masses are often reported in *atomic mass units*, a system in which the atomic mass of neutral ^{12}C is defined to be exactly 12 u.

Precise atomic masses are available in tables on the Web and are usually provided in homework problems. However, sometimes we need only an approximation of the mass of either a nucleus alone or a neutral atom. The mass number A of a nuclide gives such an approximate mass in atomic mass units. For example, the approximate mass of both the nucleus and the neutral atom for ^{197}Au is 197 u, which is close to the actual atomic mass of 196.966 552 u.

As we saw in Section 37-12,

$$1 \text{ u} = 1.660\ 538\ 86 \times 10^{-27} \text{ kg.} \quad (42-4)$$

We also saw that if the total mass of the participants in a nuclear reaction changes by an amount Δm , there is an energy release or absorption given by Eq. 37-50 ($Q = -\Delta m c^2$). As we shall now see, nuclear energies are often reported in multiples of 1 MeV. Thus, a convenient conversion between mass units and energy units is provided by Eq. 37-46:

$$c^2 = 931.494\ 013 \text{ MeV/u.} \quad (42-5)$$

Scientists and engineers working with atomic masses often prefer to report the mass of an atom by means of the atom's *mass excess* Δ , defined as

$$\Delta = M - A \quad (\text{mass excess}), \quad (42-6)$$

where M is the actual mass of the atom in atomic mass units and A is the mass number for that atom's nucleus.

Nuclear Binding Energies

The mass M of a nucleus is *less* than the total mass Σm of its individual protons and neutrons. That means that the mass energy Mc^2 of a nucleus is *less* than the total mass energy $\Sigma(mc^2)$ of its individual protons and neutrons. The difference between these two energies is called the **binding energy** of the nucleus:

$$\Delta E_{\text{be}} = \Sigma(mc^2) - Mc^2 \quad (\text{binding energy}). \quad (42-7)$$

Caution: Binding energy is not an energy that resides in the nucleus. Rather, it is a *difference* in mass energy between a nucleus and its individual nucleons: If we were able to separate a nucleus into its nucleons, we would have to transfer a total energy equal to ΔE_{be} to those particles during the separating process. Although we cannot actually tear apart a nucleus in this way, the nuclear binding energy is still a convenient measure of how well a nucleus is held together, in the sense that it measures how difficult the nucleus would be to take apart.

A better measure is the **binding energy per nucleon** ΔE_{ben} , which is the ratio of the binding energy ΔE_{be} of a nucleus to the number A of nucleons in that nucleus:

$$\Delta E_{\text{ben}} = \frac{\Delta E_{\text{be}}}{A} \quad (\text{binding energy per nucleon}). \quad (42-8)$$

We can think of the binding energy per nucleon as the average energy needed to separate a nucleus into its individual nucleons. *A greater binding energy per nucleon means a more tightly bound nucleus.*

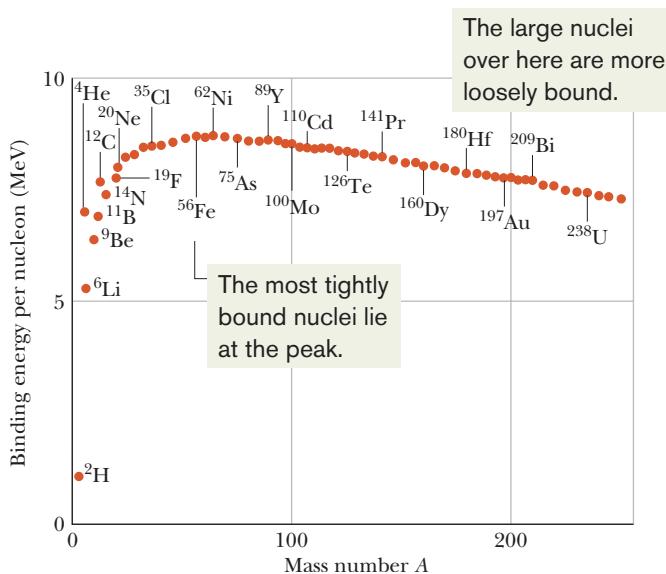


Fig. 42-7 The binding energy per nucleon for some representative nuclides. The nickel nuclide ^{62}Ni has the highest binding energy per nucleon (about 8.794 60 MeV/nucleon) of any known stable nuclide. Note that the alpha particle (^4He) has a higher binding energy per nucleon than its neighbors in the periodic table and thus is also particularly stable.

Figure 42-7 is a plot of the binding energy per nucleon ΔE_{ben} versus mass number A for a large number of nuclei. Those high on the plot are very tightly bound; that is, we would have to supply a great amount of energy per nucleon to break apart one of those nuclei. The nuclei that are lower on the plot, at the left and right sides, are less tightly bound, and less energy per nucleon would be required to break them apart.

These simple statements about Fig. 42-7 have profound consequences. The nucleons in a nucleus on the right side of the plot would be more tightly bound if that nucleus were to split into two nuclei that lie near the top of the plot. Such a process, called **fission**, occurs naturally with large (high mass number A) nuclei such as uranium, which can undergo fission spontaneously (that is, without an external cause or source of energy). The process can also occur in nuclear weapons, in which many uranium or plutonium nuclei are made to fission all at once, to create an explosion.

The nucleons in any pair of nuclei on the left side of the plot would be more tightly bound if the pair were to combine to form a single nucleus that lies near the top of the plot. Such a process, called **fusion**, occurs naturally in stars. Were this not true, the Sun would not shine and thus life could not exist on Earth.

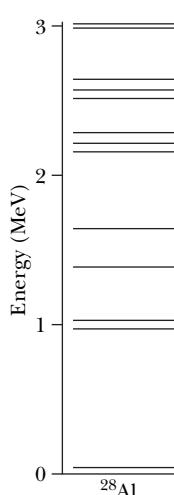


Fig. 42-8 Energy levels for the nuclide ^{28}Al , deduced from nuclear reaction experiments.

Nuclear Energy Levels

The energy of nuclei, like that of atoms, is quantized. That is, nuclei can exist only in discrete quantum states, each with a well-defined energy. Figure 42-8 shows some of these energy levels for ^{28}Al , a typical low-mass nuclide. Note that the energy scale is in millions of electron-volts, rather than the electron-volts used for atoms. When a nucleus makes a transition from one level to a level of lower energy, the emitted photon is typically in the gamma-ray region of the electromagnetic spectrum.

Nuclear Spin and Magnetism

Many nuclides have an intrinsic *nuclear angular momentum*, or spin, and an associated intrinsic *nuclear magnetic moment*. Although nuclear angular momenta are roughly of the same magnitude as the angular momenta of atomic electrons, nuclear magnetic moments are much smaller than typical atomic magnetic moments.

The Nuclear Force

The force that controls the motions of atomic electrons is the familiar electromagnetic force. To bind the nucleus together, however, there must be a strong attractive nuclear force of a totally different kind, strong enough to overcome the repulsive force between the (positively charged) nuclear protons and to bind both protons and neutrons into the tiny nuclear volume. The nuclear force must also be of short range because its influence does not extend very far beyond the nuclear “surface.”

The present view is that the nuclear force that binds neutrons and protons in the nucleus is not a fundamental force of nature but is a secondary, or “spillover,” effect of the **strong force** that binds quarks together to form neutrons and protons. In much the same way, the attractive force between certain neutral molecules is a spillover effect of the Coulomb electric force that acts within each molecule to bind it together.

Sample Problem

Binding energy per nucleon

What is the binding energy per nucleon for ^{120}Sn ?

KEY IDEAS

- We can find the binding energy per nucleon ΔE_{ben} if we first find the binding energy ΔE_{be} and then divide by the number of nucleons A in the nucleus, according to Eq. 42-8 ($\Delta E_{\text{ben}} = \Delta E_{\text{be}}/A$).
- We can find ΔE_{be} by finding the difference between the mass energy Mc^2 of the nucleus and the total mass energy $\Sigma(mc^2)$ of the individual nucleons that make up the nucleus, according to Eq. 42-7 ($\Delta E_{\text{be}} = \Sigma(mc^2) - Mc^2$).

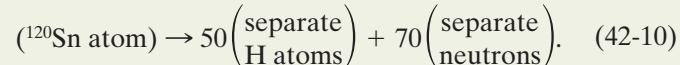
Calculations: From Table 42-1, we see that a ^{120}Sn nucleus consists of 50 protons ($Z = 50$) and 70 neutrons ($N = A - Z = 120 - 50 = 70$). Thus, we need to imagine a ^{120}Sn nucleus being separated into its 50 protons and 70 neutrons,

$$(^{120}\text{Sn nucleus}) \rightarrow 50 \left(\begin{array}{l} \text{separate} \\ \text{protons} \end{array} \right) + 70 \left(\begin{array}{l} \text{separate} \\ \text{neutrons} \end{array} \right), \quad (42-9)$$

and then compute the resulting change in mass energy.

For that computation, we need the masses of a ^{120}Sn nucleus, a proton, and a neutron. However, because the mass of a neutral atom (nucleus *plus* electrons) is much easier to measure than the mass of a bare nucleus, calculations of binding energies are traditionally done with atomic masses. Thus, let's modify Eq. 42-9 so that it has a neutral ^{120}Sn atom on the left side. To do that, we include 50 electrons on the left side (to match the 50 protons in the ^{120}Sn nucleus). We

must also add 50 electrons on the right side to balance Eq. 42-9. Those 50 electrons can be combined with the 50 protons, to form 50 neutral hydrogen atoms. We then have



From the mass column of Table 42-1, the mass M_{Sn} of a ^{120}Sn atom is 119.902 197 u and the mass m_{H} of a hydrogen atom is 1.007 825 u; the mass m_{n} of a neutron is 1.008 665 u. Thus, Eq. 42-7 yields

$$\begin{aligned} \Delta E_{\text{be}} &= \Sigma(mc^2) - Mc^2 \\ &= 50(m_{\text{H}}c^2) + 70(m_{\text{n}}c^2) - M_{\text{Sn}}c^2 \\ &= 50(1.007825 \text{ u})c^2 + 70(1.008665 \text{ u})c^2 \\ &\quad - (119.902197 \text{ u})c^2 \\ &= (1.095603 \text{ u})c^2 \\ &= (1.095603 \text{ u})(931.494013 \text{ MeV/u}) \\ &= 1020.5 \text{ MeV}, \end{aligned}$$

where Eq. 42-5 ($c^2 = 931.494013 \text{ MeV/u}$) provides an easy unit conversion. Note that using atomic masses instead of nuclear masses does not affect the result because the mass of the 50 electrons in the ^{120}Sn atom subtracts out from the mass of the electrons in the 50 hydrogen atoms.

Now Eq. 42-8 gives us the binding energy per nucleon as

$$\begin{aligned} \Delta E_{\text{ben}} &= \frac{\Delta E_{\text{be}}}{A} = \frac{1020.5 \text{ MeV}}{120} \\ &= 8.50 \text{ MeV/nucleon}. \quad (\text{Answer}) \end{aligned}$$



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Sample Problem**Density of nuclear matter**

We can think of all nuclides as made up of a neutron–proton mixture that we can call *nuclear matter*. What is the density of nuclear matter?

KEY IDEA

We can find the (average) density ρ of a nucleus by dividing its total mass by its volume.

Calculations: Let m represent the mass of a nucleon (either a proton or a neutron, because those particles have about the same mass). Then the mass of a nucleus containing A nucleons is Am . Next, we assume the nucleus is spherical with radius r . Then its volume is $\frac{4}{3}\pi r^3$, and we can write the density of the nucleus as

$$\rho = \frac{Am}{\frac{4}{3}\pi r^3}.$$

The radius r is given by Eq. 42-3 ($r = r_0 A^{1/3}$), where r_0 is 1.2 fm ($= 1.2 \times 10^{-15}$ m). Substituting for r then leads to

$$\rho = \frac{Am}{\frac{4}{3}\pi r_0^3 A} = \frac{m}{\frac{4}{3}\pi r_0^3}.$$

Note that A has canceled out; thus, this equation for density ρ applies to any nucleus that can be treated as spherical with a radius given by Eq. 42-3. Using 1.67×10^{-27} kg for the mass m of a nucleon, we then have

$$\rho = \frac{1.67 \times 10^{-27} \text{ kg}}{\frac{4}{3}\pi(1.2 \times 10^{-15} \text{ m})^3} \approx 2 \times 10^{17} \text{ kg/m}^3. \quad (\text{Answer})$$

This is about 2×10^{14} times the density of water and is the density of neutron stars, which contain only neutrons.



Additional examples, video, and practice available at WileyPLUS

42-4 Radioactive Decay

As Fig. 42-5 shows, most of the nuclides that have been identified are radioactive. A radioactive nuclide spontaneously emits a particle, transforming itself in the process into a different nuclide, occupying a different square on the nuclidic chart.

Radioactive decay provided the first evidence that the laws that govern the subatomic world are statistical. Consider, for example, a 1 mg sample of uranium metal. It contains 2.5×10^{18} atoms of the very long-lived radionuclide ^{238}U . The nuclei of these particular atoms have existed without decaying since they were created—well before the formation of our solar system. During any given second, only about 12 of the nuclei in our sample will happen to decay by emitting an alpha particle, transforming themselves into nuclei of ^{234}Th . However,



There is absolutely no way to predict whether any given nucleus in a radioactive sample will be among the small number of nuclei that decay during the next second. All have the same chance.

Although we cannot predict which nuclei in a sample will decay, we can say that if a sample contains N radioactive nuclei, then the rate ($= -dN/dt$) at which nuclei will decay is proportional to N :

$$-\frac{dN}{dt} = \lambda N, \quad (42-11)$$

in which λ , the **disintegration constant** (or **decay constant**) has a characteristic value for every radionuclide. Its SI unit is the inverse second (s^{-1}).

To find N as a function of time t , we first rearrange Eq. 42-11 as

$$\frac{dN}{N} = -\lambda dt, \quad (42-12)$$

and then integrate both sides, obtaining

$$\int_{N_0}^N \frac{dN}{N} = -\lambda \int_{t_0}^t dt,$$

or $\ln N - \ln N_0 = -\lambda(t - t_0).$ (42-13)

Here N_0 is the number of radioactive nuclei in the sample at some arbitrary initial time $t_0.$ Setting $t_0 = 0$ and rearranging Eq. 42-13 give us

$$\ln \frac{N}{N_0} = -\lambda t. \quad (42-14)$$

Taking the exponential of both sides (the exponential function is the antifunction of the natural logarithm) leads to

$$\frac{N}{N_0} = e^{-\lambda t}$$

or $N = N_0 e^{-\lambda t} \quad (\text{radioactive decay}),$ (42-15)

in which N_0 is the number of radioactive nuclei in the sample at $t = 0$ and N is the number remaining at any subsequent time $t.$ Note that lightbulbs (for one example) follow no such exponential decay law. If we life-test 1000 bulbs, we expect that they will all “decay” (that is, burn out) at more or less the same time. The decay of radionuclides follows quite a different law.

We are often more interested in the decay rate $R (= -dN/dt)$ than in N itself. Differentiating Eq. 42-15, we find

$$R = -\frac{dN}{dt} = \lambda N_0 e^{-\lambda t}$$

or $R = R_0 e^{-\lambda t} \quad (\text{radioactive decay}),$ (42-16)

an alternative form of the law of radioactive decay (Eq. 42-15). Here R_0 is the decay rate at time $t = 0$ and R is the rate at any subsequent time $t.$ We can now rewrite Eq. 42-11 in terms of the decay rate R of the sample as

$$R = \lambda N, \quad (42-17)$$

where R and the number of radioactive nuclei N that have not yet undergone decay must be evaluated at the same instant.

The total decay rate R of a sample of one or more radionuclides is called the **activity** of that sample. The SI unit for activity is the **becquerel**, named for Henri Becquerel, the discoverer of radioactivity:

$$1 \text{ becquerel} = 1 \text{ Bq} = 1 \text{ decay per second.}$$

An older unit, the **curie**, is still in common use:

$$1 \text{ curie} = 1 \text{ Ci} = 3.7 \times 10^{10} \text{ Bq.}$$

Here is an example using these units: “The activity of spent reactor fuel rod #5658 on January 15 was 3.5×10^{15} Bq ($= 9.5 \times 10^4$ Ci).” Thus, on that day 3.5×10^{15} radioactive nuclei in the rod decayed each second.

Often a radioactive sample will be placed near a detector that, for reasons of geometry or detector inefficiency, does not record all the disintegrations that occur in the sample. The reading of the detector under these circumstances is proportional to (and smaller than) the true activity of the sample. Such proportional activity measurements are reported not in becquerel units but simply in counts per unit time.

There are two common time measures of how long any given type of radionuclides lasts. One measure is the **half-life** $T_{1/2}$ of a radionuclide, which is the time at which both N and R have been reduced to one-half their initial values. The other measure is the **mean life** τ , which is the time at which both N and R have been reduced to e^{-1} of their initial values.

To relate $T_{1/2}$ to the disintegration constant λ , we put $R = \frac{1}{2}R_0$ in Eq. 42-16 and substitute $T_{1/2}$ for $t.$ We obtain

$$\frac{1}{2}R_0 = R_0 e^{-\lambda T_{1/2}}.$$

Taking the natural logarithm of both sides and solving for $T_{1/2}$, we find

$$T_{1/2} = \frac{\ln 2}{\lambda}.$$

Similarly, to relate τ to λ , we put $R = e^{-1}R_0$ in Eq. 42-16, substitute τ for t , and solve for τ , finding

$$\tau = \frac{1}{\lambda}.$$

We summarize these results with the following:

$$T_{1/2} = \frac{\ln 2}{\lambda} = \tau \ln 2. \quad (42-18)$$

Sample Problem

Finding the disintegration constant and half-life from a graph

The table that follows shows some measurements of the decay rate of a sample of ^{128}I , a radionuclide often used medically as a tracer to measure the rate at which iodine is absorbed by the thyroid gland.

Time (min)	R (counts/s)	Time (min)	R (counts/s)
4	392.2	132	10.9
36	161.4	164	4.56
68	65.5	196	1.86
100	26.8	218	1.00

Find the disintegration constant λ and the half-life $T_{1/2}$ for this radionuclide.

KEY IDEAS

The disintegration constant λ determines the exponential rate at which the decay rate R decreases with time t (as indicated by Eq. 42-16, $R = R_0 e^{-\lambda t}$). Therefore, we should be able to determine λ by plotting the measurements of R against the measurement times t . However, obtaining λ from a plot of R versus t is difficult because R decreases exponentially with t , according to Eq. 42-16. A neat solution is to transform Eq. 42-16 into a linear function of t , so that we can easily find λ . To do so, we take the natural logarithms of both sides of Eq. 42-16.

Calculations: We obtain

$$\begin{aligned} \ln R &= \ln(R_0 e^{-\lambda t}) = \ln R_0 + \ln(e^{-\lambda t}) \\ &= \ln R_0 - \lambda t. \end{aligned} \quad (42-19)$$

Because Eq. 42-19 is of the form $y = b + mx$, with b and m constants, it is a linear equation giving the quantity $\ln R$ as a

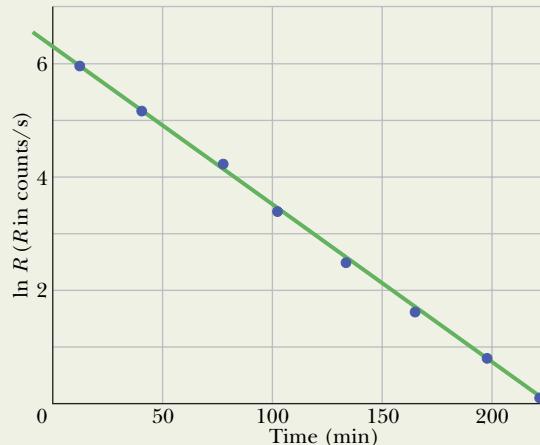


Fig. 42-9 A semilogarithmic plot of the decay of a sample of ^{128}I , based on the data in the table.

function of t . Thus, if we plot $\ln R$ (instead of R) versus t , we should get a straight line. Further, the slope of the line should be equal to $-\lambda$.

Figure 42-9 shows a plot of $\ln R$ versus time t for the given measurements. The slope of the straight line that fits through the plotted points is

$$\text{slope} = \frac{0 - 6.2}{225 \text{ min} - 0} = -0.0276 \text{ min}^{-1}.$$

$$\text{Thus, } -\lambda = -0.0276 \text{ min}^{-1}$$

$$\text{or } \lambda = 0.0276 \text{ min}^{-1} \approx 1.7 \text{ h}^{-1}. \quad (\text{Answer})$$

The time for the decay rate R to decrease by 1/2 is related to the disintegration constant λ via Eq. 42-18 ($T_{1/2} = (\ln 2)/\lambda$). From that equation, we find

$$T_{1/2} = \frac{\ln 2}{\lambda} = \frac{\ln 2}{0.0276 \text{ min}^{-1}} \approx 25 \text{ min.} \quad (\text{Answer})$$



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Sample Problem

Finding the half-life from the activity and mass

A 2.71 g sample of KCl from the chemistry stockroom is found to be radioactive, and it is decaying at a constant rate of 44.90 Bq. The decays are traced to the element potassium and in particular to the isotope ^{40}K , which constitutes 0.0117% of normal potassium. Calculate the half-life of this nuclide.

KEY IDEAS

- Because the activity R of the sample is apparently constant, we cannot find the half-life $T_{1/2}$ by plotting $\ln R$ versus time t as in the preceding sample problem. (We would just get a horizontal plot.) However, we can use the following ideas.
- We can relate the half-life $T_{1/2}$ to the disintegration constant λ via Eq. 42-18 ($T_{1/2} = (\ln 2)/\lambda$).
- We can then relate λ to the given activity R of 44.90 Bq by means of Eq. 42-17 ($R = \lambda N$), where N is the number of ^{40}K nuclei (and thus atoms) in the sample.

Calculations: Combining Eqs. 42-18 and 42-17 yields

$$T_{1/2} = \frac{N \ln 2}{R}. \quad (42-20)$$

We know that N in this equation is 0.0117% of the total number N_K of potassium atoms in the sample. We also know that N_K must equal the number N_{KCl} of molecules in the sample. We can obtain N_{KCl} from the molar mass M_{KCl} of KCl (the mass of one mole of KCl) and the given mass M_{sam} of the sample by combining Eqs. 19-2 ($n = N/N_A$) and 19-3 ($n = M_{\text{sam}}/M$) to write



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$$N_{\text{KCl}} = \left(\frac{\text{number of moles}}{\text{in sample}} \right) N_A = \frac{M_{\text{sam}}}{M_{\text{KCl}}} N_A, \quad (42-21)$$

where N_A is Avogadro's number ($6.02 \times 10^{23} \text{ mol}^{-1}$). From Appendix F, we see that the molar mass of potassium is 39.102 g/mol and the molar mass of chlorine is 35.453 g/mol; thus, the molar mass of KCl is 74.555 g/mol. Equation 42-21 then gives us

$$N_{\text{KCl}} = \frac{(2.71 \text{ g})(6.02 \times 10^{23} \text{ mol}^{-1})}{74.555 \text{ g/mol}} = 2.188 \times 10^{22}$$

as the number of KCl molecules in the sample. Thus, the total number N_K of potassium atoms is also 2.188×10^{22} , and the number of ^{40}K in the sample must be

$$\begin{aligned} N &= (0.000\,117)N_K = (0.000\,117)(2.188 \times 10^{22}) \\ &= 2.560 \times 10^{18}. \end{aligned}$$

Substituting this value for N and the given activity of 44.90 Bq ($= 44.90 \text{ s}^{-1}$) for R into Eq. 42-20 leads to

$$\begin{aligned} T_{1/2} &= \frac{(2.560 \times 10^{18}) \ln 2}{44.90 \text{ s}^{-1}} \\ &= 3.95 \times 10^{16} \text{ s} = 1.25 \times 10^9 \text{ y.} \quad (\text{Answer}) \end{aligned}$$

This half-life of ^{40}K turns out to have the same order of magnitude as the age of the universe. Thus, the activity of ^{40}K in the stockroom sample decreases *very* slowly, too slowly for us to detect during a few days of observation or even an entire lifetime. A portion of the potassium in our bodies consists of this radioisotope, which means that we are all slightly radioactive.

42-5 Alpha Decay

When a nucleus undergoes **alpha decay**, it transforms to a different nuclide by emitting an alpha particle (a helium nucleus, ^4He). For example, when uranium ^{238}U undergoes alpha decay, it transforms to thorium ^{234}Th :



This alpha decay of ^{238}U can occur spontaneously (without an external source of energy) because the total mass of the decay products ^{234}Th and ^4He is less than the mass of the original ^{238}U . Thus, the total mass energy of the decay products is less than the mass energy of the original nuclide. As defined by Eq. 37-50 ($Q = -\Delta M c^2$), in such a process the difference between the initial mass energy and the total final mass energy is called the Q of the process.

For a nuclear decay, we say that the difference in mass energy is the decay's *disintegration energy* Q . The Q for the decay in Eq. 42-22 is 4.25 MeV—that

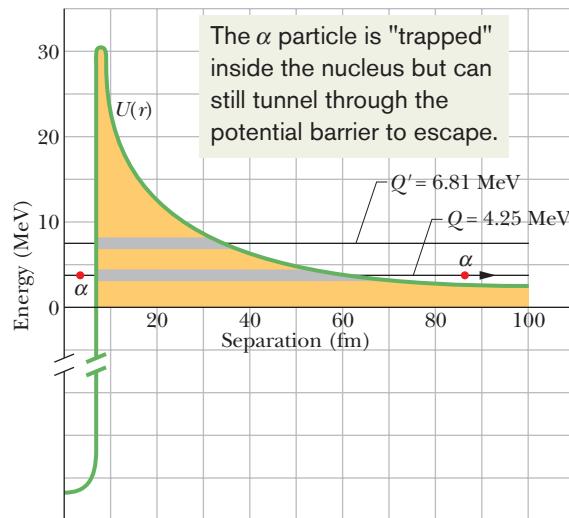


Fig. 42-10 A potential energy function for the emission of an alpha particle by ^{238}U . The horizontal black line marked $Q = 4.25 \text{ MeV}$ shows the disintegration energy for the process. The thick gray portion of this line represents separations r that are classically forbidden to the alpha particle. The alpha particle is represented by a dot, both inside this potential energy barrier (at the left) and outside it (at the right), after the particle has tunneled through. The horizontal black line marked $Q' = 6.81 \text{ MeV}$ shows the disintegration energy for the alpha decay of ^{228}U . (Both isotopes have the same potential energy function because they have the same nuclear charge.)

amount of energy is said to be released by the alpha decay of ^{238}U , with the energy transferred from mass energy to the kinetic energy of the two products.

The half-life of ^{238}U for this decay process is $4.5 \times 10^9 \text{ y}$. Why so long? If ^{238}U can decay in this way, why doesn't every ^{238}U nuclide in a sample of ^{238}U atoms simply decay at once? To answer the questions, we must examine the process of alpha decay.

We choose a model in which the alpha particle is imagined to exist (already formed) inside the nucleus before it escapes from the nucleus. Figure 42-10 shows the approximate potential energy $U(r)$ of the system consisting of the alpha particle and the residual ^{234}Th nucleus, as a function of their separation r . This energy is a combination of (1) the potential energy associated with the (attractive) strong nuclear force that acts in the nuclear interior and (2) a Coulomb potential associated with the (repulsive) electric force that acts between the two particles before and after the decay has occurred.

The horizontal black line marked $Q = 4.25 \text{ MeV}$ shows the disintegration energy for the process. If we assume that this represents the total energy of the alpha particle during the decay process, then the part of the $U(r)$ curve above this line constitutes a potential energy barrier like that in Fig. 38-15. This barrier cannot be surmounted. If the alpha particle were able to be at some separation r within the barrier, its potential energy U would exceed its total energy E . This would mean, classically, that its kinetic energy K (which equals $E - U$) would be negative, an impossible situation.

We can see now why the alpha particle is not immediately emitted from the ^{238}U nucleus. That nucleus is surrounded by an impressive potential barrier, occupying—if you think of it in three dimensions—the volume lying between two spherical shells (of radii about 8 and 60 fm). This argument is so convincing that we now change our last question and ask: Since the particle seems permanently trapped inside the nucleus by the barrier, how can the ^{238}U nucleus ever emit an alpha particle? The answer is that, as you learned in Section 38-9, there is a finite probability that a particle can tunnel through an energy barrier that is classically insurmountable. In fact, alpha decay occurs as a result of barrier tunneling.

The very long half-life of ^{238}U tells us that the barrier is apparently not very “leaky.” If we imagine that an alpha particle is already formed inside the nucleus, we can also imagine that it is rattling back and forth between the constraints of the potential barrier surrounding the nucleus. It would arrive at the inner surface of the barrier about 10^{38} times before it would succeed in tunneling through the barrier. This is about 10^{21} times per second for about 4×10^9 years (the age of Earth)! We, of course, are waiting on the outside, able to count only the alpha particles that *do* manage to escape without being able to tell what’s going on inside the nucleus.

We can test this explanation of alpha decay by examining other alpha emitters. For an extreme contrast, consider the alpha decay of another uranium isotope, ^{228}U , which has a disintegration energy Q' of 6.81 MeV, about 60% higher than that of ^{238}U . (The value of Q' is also shown as a horizontal black line in Fig. 42-10.) Recall from Section 38-9 that the transmission coefficient of a barrier is very sensitive to small changes in the total energy of the particle seeking to penetrate it. Thus, we expect alpha decay to occur more readily for this nuclide than for ^{238}U . Indeed it does. As Table 42-2 shows, its half-life is only 9.1 min! An increase in Q by a factor of only 1.6 produces a decrease in half-life (that is, in the effectiveness of the barrier) by a factor of 3×10^{14} . This is sensitivity indeed.

Table 42-2

Two Alpha Emitters Compared

Radionuclide	Q	Half-Life
^{238}U	4.25 MeV	4.5×10^9 y
^{228}U	6.81 MeV	9.1 min

Sample Problem

Q value in an alpha decay, using masses

We are given the following atomic masses:

$$\begin{array}{ll} ^{238}\text{U} & 238.050\ 79\ \text{u} \\ ^{234}\text{Th} & 234.043\ 63\ \text{u} \\ ^{237}\text{Pa} & 237.051\ 21\ \text{u} \end{array} \quad \begin{array}{ll} ^4\text{He} & 4.002\ 60\ \text{u} \\ ^1\text{H} & 1.007\ 83\ \text{u} \end{array}$$

Here Pa is the symbol for the element protactinium ($Z = 91$).

(a) Calculate the energy released during the alpha decay of ^{238}U . The decay process is



Note, incidentally, how nuclear charge is conserved in this equation: The atomic numbers of thorium (90) and helium (2) add up to the atomic number of uranium (92). The number of nucleons is also conserved: $238 = 234 + 4$.

KEY IDEA

The energy released in the decay is the disintegration energy Q , which we can calculate from the change in mass ΔM due to the ^{238}U decay.

Calculation: To do this, we use Eq. 37-50,

$$Q = M_i c^2 - M_f c^2, \quad (42-23)$$

where the initial mass M_i is that of ^{238}U and the final mass M_f is the sum of the ^{234}Th and ^4He masses. Using the atomic masses given in the problem statement, Eq. 42-23 becomes

$$\begin{aligned} Q &= (238.050\ 79\ \text{u})c^2 - (234.043\ 63\ \text{u} + 4.002\ 60\ \text{u})c^2 \\ &= (0.004\ 56\ \text{u})c^2 = (0.004\ 56\ \text{u})(931.494\ 013\ \text{MeV/u}) \\ &= 4.25\ \text{MeV}. \end{aligned} \quad (\text{Answer})$$

Note that using atomic masses instead of nuclear masses does not affect the result because the total mass of the electrons in the products subtracts out from the mass of the nucleons + electrons in the original ^{238}U .

(b) Show that ^{238}U cannot spontaneously emit a proton; that is, protons do not leak out of the nucleus in spite of the proton-proton repulsion within the nucleus.

Solution: If this happened, the decay process would be



(You should verify that both nuclear charge and the number of nucleons are conserved in this process.) Using the same Key Idea as in part (a) and proceeding as we did there, we would find that the mass of the two decay products

$$237.051\ 21\ \text{u} + 1.007\ 83\ \text{u}$$

would exceed the mass of ^{238}U by $\Delta m = 0.008\ 25\ \text{u}$, with disintegration energy

$$Q = -7.68\ \text{MeV}.$$

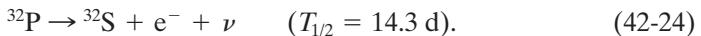
The minus sign indicates that we must *add* 7.68 MeV to a ^{238}U nucleus before it will emit a proton; it will certainly not do so spontaneously.



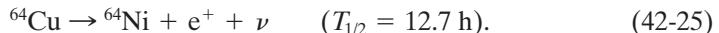
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42-6 Beta Decay

A nucleus that decays spontaneously by emitting an electron or a positron (a positively charged particle with the mass of an electron) is said to undergo **beta decay**. Like alpha decay, this is a spontaneous process, with a definite disintegration energy and half-life. Again like alpha decay, beta decay is a statistical process, governed by Eqs. 42-15 and 42-16. In *beta-minus* (β^-) decay, an electron is emitted by a nucleus, as in the decay



In *beta-plus* (β^+) decay, a positron is emitted by a nucleus, as in the decay



The symbol ν represents a **neutrino**, a neutral particle which has a very small mass, that is emitted from the nucleus along with the electron or positron during the decay process. Neutrinos interact only very weakly with matter and—for that reason—are so extremely difficult to detect that their presence long went unnoticed.*

Both charge and nucleon number are conserved in the above two processes. In the decay of Eq. 42-24, for example, we can write for charge conservation

$$(+15e) = (+16e) + (-e) + (0),$$

because ${}^{32}\text{P}$ has 15 protons, ${}^{32}\text{S}$ has 16 protons, and the neutrino ν has zero charge. Similarly, for nucleon conservation, we can write

$$(32) = (32) + (0) + (0),$$

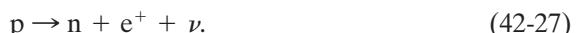
because ${}^{32}\text{P}$ and ${}^{32}\text{S}$ each have 32 nucleons and neither the electron nor the neutrino is a nucleon.

It may seem surprising that nuclei can emit electrons, positrons, and neutrinos, since we have said that nuclei are made up of neutrons and protons only. However, we saw earlier that atoms emit photons, and we certainly do not say that atoms “contain” photons. We say that the photons are created during the emission process.

It is the same with the electrons, positrons, and neutrinos emitted from nuclei during beta decay. They are created during the emission process. For beta-minus decay, a neutron transforms into a proton within the nucleus according to



For beta-plus decay, a proton transforms into a neutron via



Both of these beta-decay processes provide evidence that—as was pointed out—neutrons and protons are not truly fundamental particles. These processes show why the mass number A of a nuclide undergoing beta decay does not change; one of its constituent nucleons simply changes its character according to Eq. 42-26 or 42-27.

In both alpha decay and beta decay, the same amount of energy is released in every individual decay of a particular radionuclide. In the alpha decay of a particular radionuclide, every emitted alpha particle has the same sharply defined kinetic energy. However, in the beta-minus decay of Eq. 42-26 with electron emission, the disintegration energy Q is shared—in varying proportions—between the emitted electron and neutrino. Sometimes the electron gets nearly all the

*Beta decay also includes *electron capture*, in which a nucleus decays by absorbing one of its atomic electrons, emitting a neutrino in the process. We do not consider that process here. Also, the neutral particle emitted in the decay process of Eq. 42-24 is actually an *antineutrino*, a distinction we shall not make in this introductory treatment.

energy, sometimes the neutrino does. In every case, however, the sum of the electron's energy and the neutrino's energy gives the same value Q . A similar sharing of energy, with a sum equal to Q , occurs in beta-plus decay (Eq. 42-27).

Thus, in beta decay the energy of the emitted electrons or positrons may range from zero up to a certain maximum K_{\max} . Figure 42-11 shows the distribution of positron energies for the beta decay of ^{64}Cu (see Eq. 42-25). The maximum positron energy K_{\max} must equal the disintegration energy Q because the neutrino has approximately zero energy when the positron has K_{\max} :

$$Q = K_{\max}. \quad (42-28)$$

The Neutrino

Wolfgang Pauli first suggested the existence of neutrinos in 1930. His neutrino hypothesis not only permitted an understanding of the energy distribution of electrons or positrons in beta decay but also solved another early beta-decay puzzle involving “missing” angular momentum.

The neutrino is a truly elusive particle; the mean free path of an energetic neutrino in water has been calculated as no less than several thousand light-years. At the same time, neutrinos left over from the big bang that presumably marked the creation of the universe are the most abundant particles of physics. Billions of them pass through our bodies every second, leaving no trace.

In spite of their elusive character, neutrinos have been detected in the laboratory. This was first done in 1953 by F. Reines and C. L. Cowan, using neutrinos generated in a high-power nuclear reactor. (In 1995, Reines received a Nobel Prize for this work.) In spite of the difficulties of detection, experimental neutrino physics is now a well-developed branch of experimental physics, with avid practitioners at laboratories throughout the world.

The Sun emits neutrinos copiously from the nuclear furnace at its core, and at night these messengers from the center of the Sun come up at us from below, Earth being almost totally transparent to them. In February 1987, light from an exploding star in the Large Magellanic Cloud (a nearby galaxy) reached Earth after having traveled for 170 000 years. Enormous numbers of neutrinos were generated in this explosion, and about 10 of them were picked up by a sensitive neutrino detector in Japan; Fig. 42-12 shows a record of their passage.

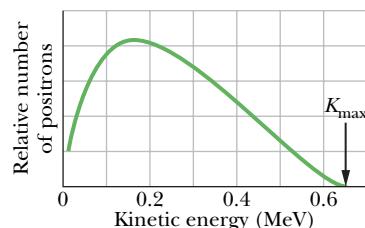


Fig. 42-11 The distribution of the kinetic energies of positrons emitted in the beta decay of ^{64}Cu . The maximum kinetic energy of the distribution (K_{\max}) is 0.653 MeV. In all ^{64}Cu decay events, this energy is shared between the positron and the neutrino, in varying proportions. The *most probable* energy for an emitted positron is about 0.15 MeV.

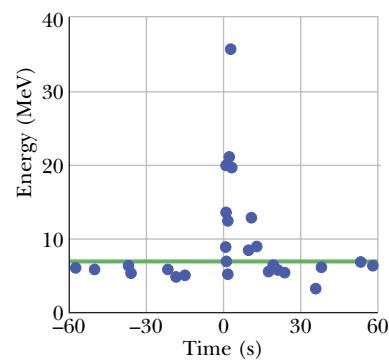


Fig. 42-12 A burst of neutrinos from the supernova SN 1987A, which occurred at (relative) time 0, stands out from the usual *background* of neutrinos. (For neutrinos, 10 is a “burst.”) The particles were detected by an elaborate detector housed deep in a mine in Japan. The supernova was visible only in the Southern Hemisphere; so the neutrinos had to penetrate Earth (a trifling barrier for them) to reach the detector.

Radioactivity and the Nuclidic Chart

We can increase the amount of information obtainable from the nuclidic chart of Fig. 42-5 by including a third axis showing the mass excess Δ expressed in the unit MeV/c^2 . The inclusion of such an axis gives Fig. 42-13, which reveals the degree of nuclear stability of the nuclides. For the low-mass nuclides, we find a “valley of the nuclides,” with the stability band of Fig. 42-5 running along its bottom. Nuclides on the proton-rich side of the valley decay into it by emitting positrons, and those on the neutron-rich side do so by emitting electrons.

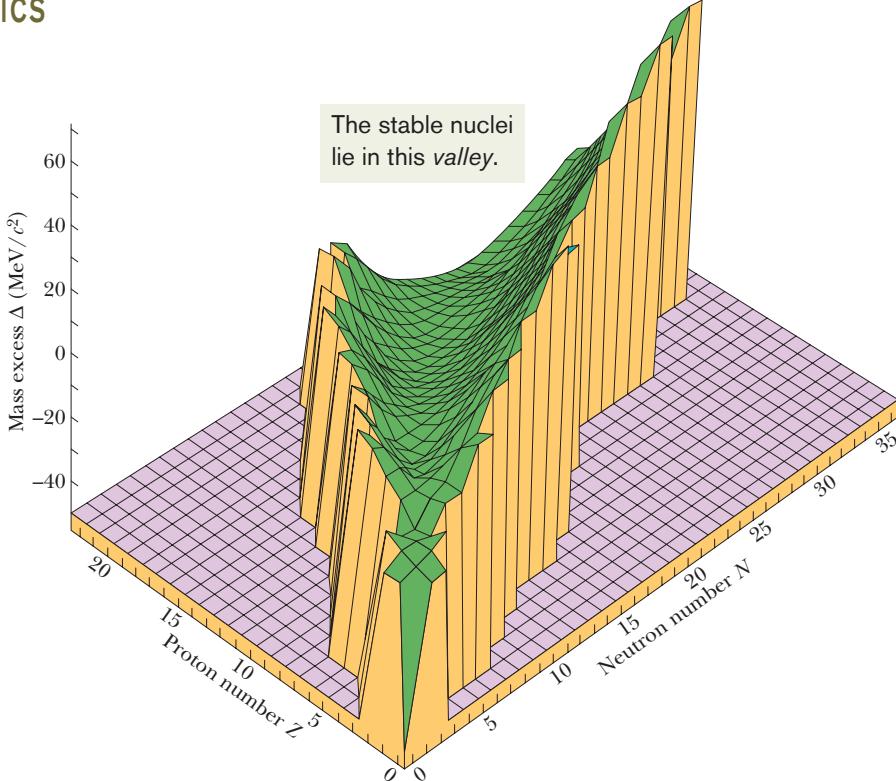


Fig. 42-13 A portion of the valley of the nuclides, showing only the nuclides of low mass. Deuterium, tritium, and helium lie at the near end of the plot, with helium at the high point. The valley stretches away from us, with the plot stopping at about $Z = 22$ and $N = 35$. Nuclides with large values of A , which would be plotted much beyond the valley, can decay into the valley by repeated alpha emissions and by fission (splitting of a nuclide).

✓ CHECKPOINT 3

^{238}U decays to ^{234}Th by the emission of an alpha particle. There follows a chain of further radioactive decays, either by alpha decay or by beta decay. Eventually a stable nuclide is reached and, after that, no further radioactive decay is possible. Which of the following stable nuclides is the end product of the ^{238}U radioactive decay chain: ^{206}Pb , ^{207}Pb , ^{208}Pb , or ^{209}Pb ? (Hint: You can decide by considering the changes in mass number A for the two types of decay.)

Sample Problem

Q value in a beta decay, using masses

Calculate the disintegration energy Q for the beta decay of ^{32}P , as described by Eq. 42-24. The needed atomic masses are 31.973 91 u for ^{32}P and 31.972 07 u for ^{32}S .

KEY IDEA

The disintegration energy Q for the beta decay is the amount by which the mass energy is changed by the decay.

Calculations: Q is given by Eq. 37-50 ($Q = -\Delta M c^2$). However, we must be careful to distinguish between nuclear masses (which we do not know) and atomic masses (which we do know). Let the boldface symbols \mathbf{m}_p and \mathbf{m}_s represent the nuclear masses of ^{32}P and ^{32}S , and let the italic symbols m_p and m_s represent their atomic masses. Then we can write the change in mass for the decay of Eq. 42-24 as

$$\Delta m = (\mathbf{m}_s + m_e) - \mathbf{m}_p,$$

in which m_e is the mass of the electron. If we add and subtract $15m_e$ on the right side of this equation, we obtain

$$\Delta m = (\mathbf{m}_s + 16m_e) - (\mathbf{m}_p + 15m_e).$$

The quantities in parentheses are the atomic masses of ^{32}S and ^{32}P ; so

$$\Delta m = m_s - m_p.$$

We thus see that if we subtract only the atomic masses, the mass of the emitted electron is automatically taken into account. (This procedure will not work for positron emission.)

The disintegration energy for the ^{32}P decay is then

$$\begin{aligned} Q &= -\Delta m c^2 \\ &= -(31.972\ 07\text{ u} - 31.973\ 91\text{ u})(931.494\ 013\text{ MeV/u}) \\ &= 1.71\text{ MeV}. \end{aligned} \quad (\text{Answer})$$

Experimentally, this calculated quantity proves to be equal to K_{\max} , the maximum energy the emitted electrons can have. Although 1.71 MeV is released every time a ^{32}P nucleus decays, in essentially every case the electron carries away less energy than this. The neutrino gets all the rest, carrying it stealthily out of the laboratory.



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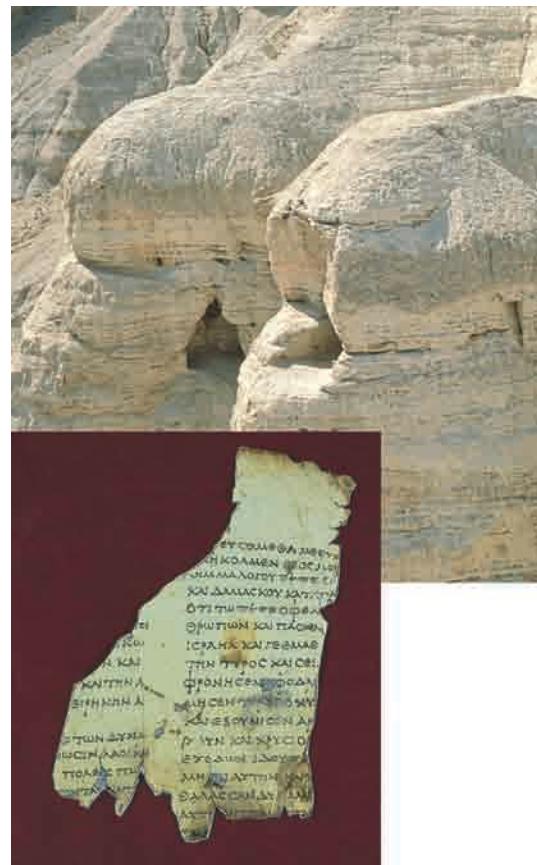
42-7 Radioactive Dating

If you know the half-life of a given radionuclide, you can in principle use the decay of that radionuclide as a clock to measure time intervals. The decay of very long-lived nuclides, for example, can be used to measure the age of rocks—that is, the time that has elapsed since they were formed. Such measurements for rocks from Earth and the Moon, and for meteorites, yield a consistent maximum age of about 4.5×10^9 y for these bodies.

The radionuclide ^{40}K , for example, decays to ^{40}Ar , a stable isotope of the noble gas argon. The half-life for this decay is 1.25×10^9 y. A measurement of the ratio of ^{40}K to ^{40}Ar , as found in the rock in question, can be used to calculate the age of that rock. Other long-lived decays, such as that of ^{235}U to ^{207}Pb (involving a number of intermediate stages of unstable nuclei), can be used to verify this calculation.

For measuring shorter time intervals, in the range of historical interest, radiocarbon dating has proved invaluable. The radionuclide ^{14}C (with $T_{1/2} = 5730$ y) is produced at a constant rate in the upper atmosphere as atmospheric nitrogen is bombarded by cosmic rays. This radiocarbon mixes with the carbon that is normally present in the atmosphere (as CO_2) so that there is about one atom of ^{14}C for every 10^{13} atoms of ordinary stable ^{12}C . Through biological activity such as photosynthesis and breathing, the atoms of atmospheric carbon trade places randomly, one atom at a time, with the atoms of carbon in every living thing, including broccoli, mushrooms, penguins, and humans. Eventually an exchange equilibrium is reached at which the carbon atoms of every living thing contain a fixed small fraction of the radioactive nuclide ^{14}C .

This equilibrium persists as long as the organism is alive. When the organism dies, the exchange with the atmosphere stops and the amount of radiocarbon trapped in the organism, since it is no longer being replenished, dwindles away with a half-life of 5730 y. By measuring the amount of radiocarbon per gram of organic matter, it is possible to measure the time that has elapsed since the organism died. Charcoal from ancient campfires, the Dead Sea scrolls (actually, the cloth used to plug the jars holding the scrolls), and many prehistoric artifacts have been dated in this way.



A fragment of the Dead Sea scrolls and the caves from which the scrolls were recovered. Top photo: (George Rockwin/Bruce Coleman, Inc.) Insert: (R. Perry/Corbis Sygma)

Sample Problem

Radioactive dating of a moon rock

In a Moon rock sample, the ratio of the number of (stable) ^{40}Ar atoms present to the number of (radioactive) ^{40}K atoms is 10.3. Assume that all the argon atoms were produced by the decay of potassium atoms, with a half-life of 1.25×10^9 y. How old is the rock?

KEY IDEAS

(1) If N_0 potassium atoms were present at the time the rock was formed by solidification from a molten form, the number of potassium atoms now remaining at the time of analysis is

$$N_{\text{K}} = N_0 e^{-\lambda t}, \quad (42-29)$$

in which t is the age of the rock. (2) For every potassium atom that decays, an argon atom is produced. Thus, the number of argon atoms present at the time of the analysis is

$$N_{\text{Ar}} = N_0 - N_{\text{K}}. \quad (42-30)$$

Calculations: We cannot measure N_0 ; so let's eliminate it from Eqs. 42-29 and 42-30. We find, after some algebra, that

$$\lambda t = \ln \left(1 + \frac{N_{\text{Ar}}}{N_{\text{K}}} \right), \quad (42-31)$$

in which $N_{\text{Ar}}/N_{\text{K}}$ can be measured. Solving for t and using Eq. 42-18 to replace λ with $(\ln 2)/T_{1/2}$ yield

$$\begin{aligned} t &= \frac{T_{1/2} \ln(1 + N_{\text{Ar}}/N_{\text{K}})}{\ln 2} \\ &= \frac{(1.25 \times 10^9 \text{ y})[\ln(1 + 10.3)]}{\ln 2} \\ &= 4.37 \times 10^9 \text{ y.} \end{aligned} \quad (\text{Answer})$$

Lesser ages may be found for other lunar or terrestrial rock samples, but no substantially greater ones. Thus, the oldest rocks were formed soon after the solar system formed, and the solar system must be about 4 billion years old.



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42-8 Measuring Radiation Dosage

The effect of radiation such as gamma rays, electrons, and alpha particles on living tissue (particularly our own) is a matter of public interest. Such radiation is found in nature in cosmic rays (from astronomical sources) and in the emissions by radioactive elements in Earth's crust. Radiation associated with some human activities, such as using x rays and radionuclides in medicine and in industry, also contributes.

Our task here is not to explore the various sources of radiation but simply to describe the units in which the properties and effects of such radiations are expressed. We have already discussed the *activity* of a radioactive source. There are two remaining quantities of interest.

- 1. Absorbed Dose.** This is a measure of the radiation dose (as energy per unit mass) actually absorbed by a specific object, such as a patient's hand or chest. Its SI unit is the **gray** (Gy). An older unit, the **rad** (from **radiation absorbed dose**) is still in common use. The terms are defined and related as follows:

$$1 \text{ Gy} = 1 \text{ J/kg} = 100 \text{ rad.} \quad (42-32)$$

A typical dose-related statement is: "A whole-body, short-term gamma-ray dose of 3 Gy (= 300 rad) will cause death in 50% of the population exposed to it." Thankfully, our present average absorbed dose per year, from sources of both natural and human origin, is only about 2 mGy (= 0.2 rad).

- 2. Dose Equivalent.** Although different types of radiation (gamma rays and neutrons, say) may deliver the same amount of energy to the body, they do not have the same biological effect. The dose equivalent allows us to express the biological effect by multiplying the absorbed dose (in grays or rads) by a numerical **RBE** factor (from **relative biological effectiveness**). For x rays and electrons, for example, RBE = 1; for slow neutrons, RBE = 5; for alpha particles, RBE = 10; and so on. Personnel-monitoring devices such as film badges register the dose equivalent.

The SI unit of dose equivalent is the **sievert** (Sv). An earlier unit, the **rem**, is still in common use. Their relationship is

$$1 \text{ Sv} = 100 \text{ rem.} \quad (42-33)$$

An example of the correct use of these terms is: "The recommendation of the National Council on Radiation Protection is that no individual who is (nonoccupationally) exposed to radiation should receive a dose equivalent greater than 5 mSv (= 0.5 rem) in any one year." This includes radiation of all kinds; of course the appropriate RBE factor must be used for each kind.

42-9 Nuclear Models

Nuclei are more complicated than atoms. For atoms, the basic force law (Coulomb's law) is simple in form and there is a natural force center, the nucleus. For nuclei, the force law is complicated and cannot, in fact, be written down explicitly in full detail. Furthermore, the nucleus—a jumble of protons and neutrons—has no natural force center to simplify the calculations.

In the absence of a comprehensive nuclear *theory*, we turn to the construction of nuclear *models*. A nuclear model is simply a way of looking at the nucleus that gives a physical insight into as wide a range of its properties as possible. The usefulness of a model is tested by its ability to provide predictions that can be verified experimentally in the laboratory.

Two models of the nucleus have proved useful. Although based on assumptions that seem flatly to exclude each other, each accounts very well for a selected group

of nuclear properties. After describing them separately, we shall see how these two models may be combined to form a single coherent picture of the atomic nucleus.

The Collective Model

In the *collective model*, formulated by Niels Bohr, the nucleons, moving around within the nucleus at random, are imagined to interact strongly with each other, like the molecules in a drop of liquid. A given nucleon collides frequently with other nucleons in the nuclear interior, its mean free path as it moves about being substantially less than the nuclear radius.

The collective model permits us to correlate many facts about nuclear masses and binding energies; it is useful (as you will see later) in explaining nuclear fission. It is also useful for understanding a large class of nuclear reactions.

Consider, for example, a generalized nuclear reaction of the form



We imagine that projectile a enters target nucleus X , forming a **compound nucleus** C and conveying to it a certain amount of excitation energy. The projectile, perhaps a neutron, is at once caught up by the random motions that characterize the nuclear interior. It quickly loses its identity—so to speak—and the excitation energy it carried into the nucleus is quickly shared with all the other nucleons in C .

The quasi-stable state represented by C in Eq. 42-34 may have a mean life of 10^{-16} s before it decays to Y and b . By nuclear standards, this is a very long time, being about one million times longer than the time required for a nucleon with a few million electron-volts of energy to travel across a nucleus.

The central feature of this compound-nucleus concept is that the formation of the compound nucleus and its eventual decay are totally independent events. At the time of its decay, the compound nucleus has “forgotten” how it was formed. Hence, its mode of decay is not influenced by its mode of formation. As an example, Fig. 42-14 shows three possible ways in which the compound nucleus ^{20}Ne might be formed and three in which it might decay. Any of the three formation modes can lead to any of the three decay modes.

The Independent Particle Model

In the collective model, we assume that the nucleons move around at random and bump into one another frequently. The *independent particle model*, however, is based on just the opposite assumption—namely, that each nucleon remains in a well-defined quantum state within the nucleus and makes hardly any collisions at all! The nucleus, unlike the atom, has no fixed center of charge; we assume in this model that each nucleon moves in a potential well that is determined by the smeared-out (time-averaged) motions of all the other nucleons.

A nucleon in a nucleus, like an electron in an atom, has a set of quantum numbers that defines its state of motion. Also, nucleons obey the Pauli exclusion principle, just as electrons do; that is, no two nucleons in a nucleus may occupy the same quantum state at the same time. In this regard, the neutrons and the protons are treated separately, each particle type with its own set of quantum states.

The fact that nucleons obey the Pauli exclusion principle helps us to understand the relative stability of nucleon states. If two nucleons within the nucleus are to collide, the energy of each of them after the collision must correspond to the energy of an *unoccupied* state. If no such state is available, the collision simply cannot occur. Thus, any given nucleon experiencing repeated “frustrated collision opportunities” will maintain its state of motion long enough to give meaning to the statement that it exists in a quantum state with a well-defined energy.

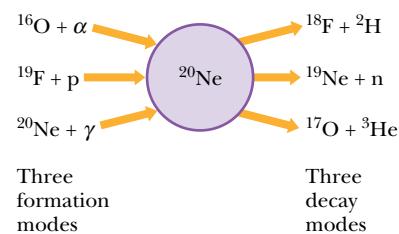


Fig. 42-14 The formation modes and the decay modes of the compound nucleus ^{20}Ne .

In the atomic realm, the repetitions of physical and chemical properties that we find in the periodic table are associated with a property of atomic electrons—namely, they arrange themselves in shells that have a special stability when fully occupied. We can take the atomic numbers of the noble gases,

$$2, 10, 18, 36, 54, 86, \dots,$$

as *magic electron numbers* that mark the completion (or closure) of such shells.

Nuclei also show such closed-shell effects, associated with certain **magic nucleon numbers**:

$$2, 8, 20, 28, 50, 82, 126, \dots$$

Any nuclide whose proton number Z or neutron number N has one of these values turns out to have a special stability that may be made apparent in a variety of ways.

Examples of “magic” nuclides are ^{18}O ($Z = 8$), ^{40}Ca ($Z = 20, N = 20$), ^{92}Mo ($N = 50$), and ^{208}Pb ($Z = 82, N = 126$). Both ^{40}Ca and ^{208}Pb are said to be “doubly magic” because they contain both filled shells of protons *and* filled shells of neutrons.

The magic number 2 shows up in the exceptional stability of the alpha particle (^4He), which, with $Z = N = 2$, is doubly magic. For example, on the binding energy curve of Fig. 42-7, the binding energy per nucleon for this nuclide stands well above those of its periodic-table neighbors hydrogen, lithium, and beryllium. The neutrons and protons making up the alpha particle are so tightly bound to one another, in fact, that it is impossible to add another proton or neutron to it; there is no stable nuclide with $A = 5$.

The central idea of a closed shell is that a single particle outside a closed shell can be relatively easily removed, but considerably more energy must be expended to remove a particle from the shell itself. The sodium atom, for example, has one (valence) electron outside a closed electron shell. Only about 5 eV is required to strip the valence electron away from a sodium atom; however, to remove a *second* electron (which must be plucked out of a closed shell) requires 22 eV. As a nuclear case, consider ^{121}Sb ($Z = 51$), which contains a single proton outside a closed shell of 50 protons. To remove this lone proton requires 5.8 MeV; to remove a *second* proton, however, requires an energy of 11 MeV. There is much additional experimental evidence that the nucleons in a nucleus form closed shells and that these shells exhibit stable properties.

We have seen that quantum theory can account beautifully for the magic electron numbers—that is, for the populations of the subshells into which atomic electrons are grouped. It turns out that, under certain assumptions, quantum theory can account equally well for the magic nucleon numbers! The 1963 Nobel Prize in physics was, in fact, awarded to Maria Mayer and Hans Jensen “for their discoveries concerning nuclear shell structure.”

A Combined Model

Consider a nucleus in which a small number of neutrons (or protons) exist outside a core of closed shells that contains magic numbers of neutrons or protons. The outside nucleons occupy quantized states in a potential well established by the central core, thus preserving the central feature of the independent-particle model. These outside nucleons also interact with the core, deforming it and setting up “tidal wave” motions of rotation or vibration within it. These collective motions of the core preserve the central feature of the collective model. Such a model of nuclear structure thus succeeds in combining the seemingly irreconcilable points of view of the collective and independent-particle models. It has been remarkably successful in explaining observed nuclear properties.

Sample Problem

Lifetime of a compound nucleus made by neutron capture

Consider the neutron capture reaction



in which a compound nucleus (^{110}Ag) is formed. Figure 42-15 shows the relative rate at which such events take place, plotted against the energy of the incoming neutron. Find the mean lifetime of this compound nucleus by using the uncertainty principle in the form

$$\Delta E \cdot \Delta t \approx \hbar. \quad (42-36)$$

Here ΔE is a measure of the uncertainty with which the energy of a state can be defined. The quantity Δt is a measure of the time available to measure this energy. In fact, here Δt is just t_{avg} , the average life of the compound nucleus before it decays to its ground state.

Reasoning: We see that the relative reaction rate peaks sharply at a neutron energy of about 5.2 eV. This suggests that we are dealing with a single excited energy level of the compound nucleus ^{110}Ag . When the available energy (of the incoming neutron) just matches the energy of this level above the ^{110}Ag ground state, we have “resonance” and the reaction of Eq. 42-35 really “goes.”

However, the resonance peak is not infinitely sharp but has an approximate half-width (ΔE in the figure) of about 0.20 eV. We can account for this resonance-peak width by saying that the excited level is not sharply defined in energy but has an energy uncertainty ΔE of about 0.20 eV.

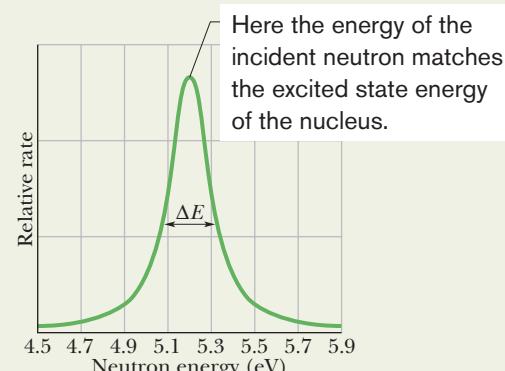


Fig. 42-15 A plot of the relative number of reaction events of the type described by Eq. 42-35 as a function of the energy of the incident neutron. The half-width ΔE of the resonance peak is about 0.20 eV.

Calculation: Substituting that uncertainty of 0.20 eV into Eq. 42-36 gives us

$$\Delta t = t_{\text{avg}} \approx \frac{\hbar}{\Delta E} \approx \frac{(4.14 \times 10^{-15} \text{ eV} \cdot \text{s})/2\pi}{0.20 \text{ eV}} \approx 3 \times 10^{-15} \text{ s.} \quad (\text{Answer})$$

This is several hundred times greater than the time a 0.20 eV neutron takes to cross the diameter of a ^{109}Ag nucleus. Therefore, the neutron is spending this time of 3×10^{-15} s *as part of the nucleus*.



Additional examples, video, and practice available at WileyPLUS

REVIEW & SUMMARY

The Nuclides Approximately 2000 nuclides are known to exist. Each is characterized by an **atomic number** Z (the number of protons), a **neutron number** N , and a **mass number** A (the total number of nucleons—protons and neutrons). Thus, $A = Z + N$. Nuclides with the same atomic number but different neutron numbers are **isotopes** of one another. Nuclei have a mean radius r given by

$$r = r_0 A^{1/3}, \quad (42-3)$$

where $r_0 \approx 1.2$ fm.

Mass and Binding Energy Atomic masses are often reported in terms of *mass excess*

$$\Delta = M - A \quad (\text{mass excess}), \quad (42-6)$$

where M is the actual mass of an atom in atomic mass units and A is the mass number for that atom's nucleus. The **binding energy** of a nucleus is the difference

$$\Delta E_{\text{be}} = \Sigma(mc^2) - Mc^2 \quad (\text{binding energy}), \quad (42-7)$$

where $\Sigma(mc^2)$ is the total mass energy of the *individual* protons and neutrons. The **binding energy per nucleon** is

$$\Delta E_{\text{ben}} = \frac{\Delta E_{\text{be}}}{A} \quad (\text{binding energy per nucleon}). \quad (42-8)$$

Mass-Energy Exchanges The energy equivalent of one mass unit (u) is 931.494 013 MeV. The binding energy curve shows that middle-mass nuclides are the most stable and that energy can be released both by fission of high-mass nuclei and by fusion of low-mass nuclei.

The Nuclear Force Nuclei are held together by an attractive force acting among the nucleons, part of the **strong force** acting between the quarks that make up the nucleons.

Radioactive Decay Most known nuclides are radioactive; they spontaneously decay at a rate R ($= -dN/dt$) that is proportional to the number N of radioactive atoms present, the

proportionality constant being the **disintegration constant** λ . This leads to the law of exponential decay:

$$N = N_0 e^{-\lambda t}, \quad R = \lambda N = R_0 e^{-\lambda t}$$

(radioactive decay). (42-15, 42-17, 42-16)

The **half-life** $T_{1/2} = (\ln 2)/\lambda$ of a radioactive nuclide is the time required for the decay rate R (or the number N) in a sample to drop to half its initial value.

Alpha Decay Some nuclides decay by emitting an alpha particle (a helium nucleus, ${}^4\text{He}$). Such decay is inhibited by a potential energy barrier that cannot be penetrated according to classical physics but is subject to tunneling according to quantum physics. The barrier penetrability, and thus the half-life for alpha decay, is very sensitive to the energy of the emitted alpha particle.

Beta Decay In **beta decay** either an electron or a positron is emitted by a nucleus, along with a neutrino. The emitted particles share the available disintegration energy. The electrons and positrons emitted in beta decay have a continuous spectrum of energies from near zero up to a limit $K_{\max} (= Q = -\Delta m c^2)$.

Radioactive Dating Naturally occurring radioactive nuclides provide a means for estimating the dates of historic and prehistoric events. For example, the ages of organic materials can often be

found by measuring their ${}^{14}\text{C}$ content; rock samples can be dated using the radioactive isotope ${}^{40}\text{K}$.

Radiation Dosage Three units are used to describe exposure to ionizing radiation. The **becquerel** (1 Bq = 1 decay per second) measures the **activity** of a source. The amount of energy actually absorbed is measured in **grays**, with 1 Gy corresponding to 1 J/kg. The estimated biological effect of the absorbed energy is measured in **sieverts**; a dose equivalent of 1 Sv causes the same biological effect regardless of the radiation type by which it was acquired.

Nuclear Models The **collective** model of nuclear structure assumes that nucleons collide constantly with one another and that relatively long-lived **compound nuclei** are formed when a projectile is captured. The formation and eventual decay of a compound nucleus are totally independent events.

The **independent particle** model of nuclear structure assumes that each nucleon moves, essentially without collisions, in a quantized state within the nucleus. The model predicts nucleon levels and **magic nucleon numbers** (2, 8, 20, 28, 50, 82, and 126) associated with closed shells of nucleons; nuclides with any of these numbers of neutrons or protons are particularly stable.

The **combined** model, in which extra nucleons occupy quantized states outside a central core of closed shells, is highly successful in predicting many nuclear properties.

** View All Solutions Here **

Q U E S T I O N S

1 The radionuclide ${}^{196}\text{Ir}$ decays by emitting an electron. (a) Into which square in Fig. 42-6 is it transformed? (b) Do further decays then occur?

2 Is the mass excess of an alpha particle (use a straightedge on Fig. 42-13) greater than or less than the particle's total binding energy (use the binding energy per nucleon from Fig. 42-7)?

3 At $t = 0$, a sample of radionuclide A has the same decay rate as a sample of radionuclide B has at $t = 30$ min. The disintegration constants are λ_A and λ_B , with $\lambda_A < \lambda_B$. Will the two samples ever have (simultaneously) the same decay rate? (*Hint:* Sketch a graph of their activities.)

4 A certain nuclide is said to be particularly stable. Does its binding energy per nucleon lie slightly above or slightly below the binding energy curve of Fig. 42-7?

5 Suppose the alpha particle in a Rutherford scattering experiment is replaced with a proton of the same initial kinetic energy and also headed directly toward the nucleus of the gold atom. (a) Will the distance from the center of the nucleus at which the proton stops be greater than, less than, or the same as that of the alpha particle? (b) If, instead, we switch the target to a nucleus with a larger value of Z , is the stopping distance of the alpha particle greater than, less than, or the same as with the gold target?

6 Figure 42-16 gives the activities of three radioactive samples versus time. Rank the samples according to their (a) half-life and (b) disintegration constant, greatest first. (*Hint:* For (a), use a straightedge on the graph.)

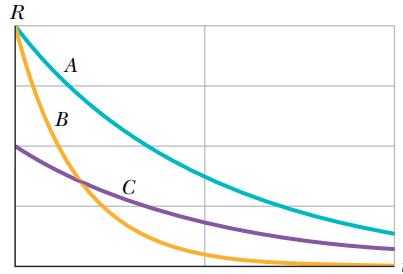


Fig. 42-16 Question 6.

7 The nuclide ${}^{244}\text{Pu}$ ($Z = 94$) is an alpha-emitter. Into which of the following nuclides does it decay: ${}^{240}\text{Np}$ ($Z = 93$), ${}^{240}\text{U}$ ($Z = 92$), ${}^{248}\text{Cm}$ ($Z = 96$), or ${}^{244}\text{Am}$ ($Z = 95$)?

8 The radionuclide ${}^{49}\text{Sc}$ has a half-life of 57.0 min. At $t = 0$, the counting rate of a sample of it is 6000 counts/min above the general background activity, which is 30 counts/min. Without computation, determine whether the counting rate of the sample will be about equal to the background rate in 3 h, 7 h, 10 h, or a time much longer than 10 h.

9 At $t = 0$ we begin to observe two identical radioactive nuclei that have a half-life of 5 min. At $t = 1$ min, one of the nuclei decays. Does that event increase or decrease the chance that the second nucleus will decay in the next 4 min, or is there no effect on the second nucleus? (Are the events cause and effect, or random?)

10 Figure 42-17 shows the curve for the binding energy per nucleon ΔE_{ben} versus mass number A . Three isotopes are indicated.

** View All Solutions Here **

PROBLEMS

1189

Rank them according to the energy required to remove a nucleon from the isotope, greatest first.

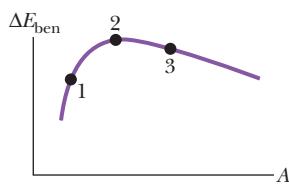


Fig. 42-17 Question 10.

- 11** At $t = 0$, a sample of radionuclide A has twice the decay rate as a sample of radionuclide B . The disintegration constants are λ_A and λ_B , with $\lambda_A > \lambda_B$. Will the two samples ever have (simultaneously) the same decay rate?

- 12** Figure 42-18 is a plot of mass number A versus charge number Z . The location of a certain nucleus is represented by a dot. Which

of the arrows extending from the dot would best represent the transition were the nucleus to undergo (a) a β^- decay and (b) an α decay?

- 13** (a) Which of the following nuclides are magic: ^{122}Sn , ^{132}Sn , ^{98}Cd , ^{198}Au , ^{208}Pb ? (b) Which, if any, are doubly magic?

- 14** If the mass of a radioactive sample is doubled, do (a) the activity of the sample and (b) the disintegration constant of the sample increase, decrease, or remain the same?

- 15** The magic nucleon numbers for nuclei are given in Section 42-9 as 2, 8, 20, 28, 50, 82, and 126. Are nuclides magic (that is, especially stable) when (a) only the mass number A , (b) only the atomic number Z , (c) only the neutron number N , or (d) either Z or N (or both) is equal to one of these numbers? Pick all correct phrases.

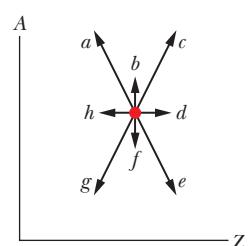


Fig. 42-18
Question 12.

P R O B L E M S



Tutoring problem available (at instructor's discretion) in *WileyPLUS* and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>



sec. 42-2 Discovering the Nucleus

- 1** A ^7Li nucleus with a kinetic energy of 3.00 MeV is sent toward a ^{232}Th nucleus. What is the least center-to-center separation between the two nuclei, assuming that the (more massive) ^{232}Th nucleus does not move?

- 2** Calculate the distance of closest approach for a head-on collision between a 5.30 MeV alpha particle and a copper nucleus.

- 3** A 10.2 MeV Li nucleus is shot directly at the center of a Ds nucleus. At what center-to-center distance does the Li momentarily stop, assuming the Ds does not move?

- 4** In a Rutherford scattering experiment, assume that an incident alpha particle (radius 1.80 fm) is headed directly toward a target gold nucleus (radius 6.23 fm). What energy must the alpha particle have to just barely “touch” the gold nucleus?

- 5** When an alpha particle collides elastically with a nucleus, the nucleus recoils. Suppose a 5.00 MeV alpha particle has a head-on elastic collision with a gold nucleus that is initially at rest. What is the kinetic energy of (a) the recoiling nucleus and (b) the rebounding alpha particle?

sec. 42-3 Some Nuclear Properties

- 6** (a) Show that the mass M of an atom is given approximately by $M_{\text{app}} = Am_p$, where A is the mass number and m_p is the proton mass. For (b) ^1H , (c) ^{31}P , (d) ^{120}Sn , (e) ^{197}Au , and (f) ^{239}Pu , use Table 42-1 to find the percentage deviation between M_{app} and M :

$$\text{percentage deviation} = \frac{M_{\text{app}} - M}{M} \cdot 100.$$

- (g) Is a value of M_{app} accurate enough to be used in a calculation of a nuclear binding energy?

- 7** What is the nuclear mass density ρ_m of (a) the fairly low-mass nuclide ^{55}Mn and (b) the fairly high-mass nuclide ^{209}Bi ? (c) Compare the two answers, with an explanation. What is the nuclear charge density ρ_q of (d) ^{55}Mn and (e) ^{209}Bi ? (f) Compare the two answers, with an explanation.

- 8** The strong neutron excess (defined as $N - Z$) of high-mass nuclei is illustrated by noting that most high-mass nuclides could never fission into two stable nuclei without neutrons being left over. For example, consider the spontaneous fission of a ^{235}U nucleus into two stable *daughter nuclei* with atomic numbers 39 and 53. From Appendix F, determine the name of the (a) first and (b) second daughter nucleus. From Fig. 42-5, approximately how many neutrons are in the (c) first and (d) second? (e) Approximately how many neutrons are left over?

- 9** The nuclide ^{14}C contains (a) how many protons and (b) how many neutrons?

- 10** What is the mass excess Δ_1 of ^1H (actual mass is 1.007 825 u) in (a) atomic mass units and (b) MeV/c^2 ? What is the mass excess Δ_n of a neutron (actual mass is 1.008 665 u) in (c) atomic mass units and (d) MeV/c^2 ? What is the mass excess Δ_{120} of ^{120}Sn (actual mass is 119.902 197 u) in (e) atomic mass units and (f) MeV/c^2 ?

- 11** **SSM** Nuclear radii may be measured by scattering high-energy (high speed) electrons from nuclei. (a) What is the de Broglie wavelength for 200 MeV electrons? (b) Are these electrons suitable probes for this purpose?

- 12** The electric potential energy of a uniform sphere of charge q and radius r is given by

$$U = \frac{3q^2}{20\pi\epsilon_0 r}.$$

1190 CHAPTER 42 NUCLEAR PHYSICS

(a) Does the energy represent a tendency for the sphere to bind together or blow apart? The nuclide ^{239}Pu is spherical with radius 6.64 fm. For this nuclide, what are (b) the electric potential energy U according to the equation, (c) the electric potential energy per proton, and (d) the electric potential energy per nucleon? The binding energy per nucleon is 7.56 MeV. (e) Why is the nuclide bound so well when the answers to (c) and (d) are large and positive?

•13 A neutron star is a stellar object whose density is about that of nuclear matter, $2 \times 10^{17} \text{ kg/m}^3$. Suppose that the Sun were to collapse and become such a star without losing any of its present mass. What would be its radius?

•14 GO What is the binding energy per nucleon of the americium isotope ^{244}Am ? Here are some atomic masses and the neutron mass.

^{244}Am	244.064 279 u	^1H	1.007 825 u
n	1.008 665 u		

•15 (a) Show that the energy associated with the strong force between nucleons in a nucleus is proportional to A , the mass number of the nucleus in question. (b) Show that the energy associated with the Coulomb force between protons in a nucleus is proportional to $Z(Z - 1)$. (c) Show that, as we move to larger and larger nuclei (see Fig. 42-5), the importance of the Coulomb force increases more rapidly than does that of the strong force.

•16 GO What is the binding energy per nucleon of the europium isotope ^{152}Eu ? Here are some atomic masses and the neutron mass.

^{152}Eu	151.921 742 u	^1H	1.007 825 u
n	1.008 665 u		

•17 Because the neutron has no charge, its mass must be found in some way other than by using a mass spectrometer. When a neutron and a proton meet (assume both to be almost stationary), they combine and form a deuteron, emitting a gamma ray whose energy is 2.2233 MeV. The masses of the proton and the deuteron are 1.007 276 467 u and 2.013 553 212 u, respectively. Find the mass of the neutron from these data.

•18 GO What is the binding energy per nucleon of the rutherfordium isotope ^{259}Rf ? Here are some atomic masses and the neutron mass.

^{259}Rf	259.105 63 u	^1H	1.007 825 u
n	1.008 665 u		

•19 A periodic table might list the average atomic mass of magnesium as being 24.312 u, which is the result of *weighting* the atomic masses of the magnesium isotopes according to their natural abundances on Earth. The three isotopes and their masses are ^{24}Mg (23.985 04 u), ^{25}Mg (24.985 84 u), and ^{26}Mg (25.982 59 u). The natural abundance of ^{24}Mg is 78.99% by mass (that is, 78.99% of the mass of a naturally occurring sample of magnesium is due to the presence of ^{24}Mg). What is the abundance of (a) ^{25}Mg and (b) ^{26}Mg ?

•20 What is the binding energy per nucleon of ^{262}Bh ? The mass of the atom is 262.1231 u.

•21 SSM WWW (a) Show that the total binding energy E_{be} of a given nuclide is

$$E_{\text{be}} = Z\Delta_H + N\Delta_n - \Delta,$$

where Δ_H is the mass excess of ^1H , Δ_n is the mass excess of a neutron,

and Δ is the mass excess of the given nuclide. (b) Using this method, calculate the binding energy per nucleon for ^{197}Au . Compare your result with the value listed in Table 42-1. The needed mass excesses, rounded to three significant figures, are $\Delta_H = +7.29 \text{ MeV}$, $\Delta_n = +8.07 \text{ MeV}$, and $\Delta_{197} = -31.2 \text{ MeV}$. Note the economy of calculation that results when mass excesses are used in place of the actual masses.

•22 GO An α particle (^4He nucleus) is to be taken apart in the following steps. Give the energy (work) required for each step: (a) remove a proton, (b) remove a neutron, and (c) separate the remaining proton and neutron. For an α particle, what are (d) the total binding energy and (e) the binding energy per nucleon? (f) Does either match an answer to (a), (b), or (c)? Here are some atomic masses and the neutron mass.

^4He	4.002 60 u	^2H	2.014 10 u
^3H	3.016 05 u	^1H	1.007 83 u
n	1.008 67 u		

•23 SSM Verify the binding energy per nucleon given in Table 42-1 for ^{239}Pu . The mass of the atom is 239.052 16 u.

•24 A penny has a mass of 3.0 g. Calculate the energy that would be required to separate all the neutrons and protons in this coin from one another. For simplicity, assume that the penny is made entirely of ^{63}Cu atoms (of mass 62.929 60 u). The masses of the proton-plus-electron and the neutron are 1.007 83 u and 1.008 66 u, respectively.

sec. 42-4 Radioactive Decay

•25 Cancer cells are more vulnerable to x and gamma radiation than are healthy cells. In the past, the standard source for radiation therapy was radioactive ^{60}Co , which decays, with a half-life of 5.27 y, into an excited nuclear state of ^{60}Ni . That nickel isotope then immediately emits two gamma-ray photons, each with an approximate energy of 1.2 MeV. How many radioactive ^{60}Co nuclei are present in a 6000 Ci source of the type used in hospitals? (Energetic particles from linear accelerators are now used in radiation therapy.)

•26 The half-life of a radioactive isotope is 140 d. How many days would it take for the decay rate of a sample of this isotope to fall to one-fourth of its initial value?

•27 A radioactive nuclide has a half-life of 30.0 y. What fraction of an initially pure sample of this nuclide will remain undecayed at the end of (a) 60.0 y and (b) 90.0 y?

•28 The plutonium isotope ^{239}Pu is produced as a by-product in nuclear reactors and hence is accumulating in our environment. It is radioactive, decaying with a half-life of 2.41×10^4 y. (a) How many nuclei of Pu constitute a chemically lethal dose of 2.00 mg? (b) What is the decay rate of this amount?

•29 SSM WWW A radioactive isotope of mercury, ^{197}Hg , decays to gold, ^{197}Au , with a disintegration constant of 0.0108 h^{-1} . (a) Calculate the half-life of the ^{197}Hg . What fraction of a sample will remain at the end of (b) three half-lives and (c) 10.0 days?

•30 The half-life of a particular radioactive isotope is 6.5 h. If there are initially 48×10^{19} atoms of this isotope, how many remain at the end of 26 h?

•31 Consider an initially pure 3.4 g sample of ^{67}Ga , an isotope that has a half-life of 78 h. (a) What is its initial decay rate? (b) What is its decay rate 48 h later?

••32 When aboveground nuclear tests were conducted, the explosions shot radioactive dust into the upper atmosphere. Global air circulations then spread the dust worldwide before it settled out on ground and water. One such test was conducted in October 1976. What fraction of the ^{90}Sr produced by that explosion still existed in October 2006? The half-life of ^{90}Sr is 29 y.

••33 The air in some caves includes a significant amount of radon gas, which can lead to lung cancer if breathed over a prolonged time. In British caves, the air in the cave with the greatest amount of the gas has an activity per volume of $1.55 \times 10^5 \text{ Bq/m}^3$. Suppose that you spend two full days exploring (and sleeping in) that cave. Approximately how many ^{222}Rn atoms would you take in and out of your lungs during your two-day stay? The radionuclide ^{222}Rn in radon gas has a half-life of 3.82 days. You need to estimate your lung capacity and average breathing rate.

••34 Calculate the mass of a sample of (initially pure) ^{40}K that has an initial decay rate of 1.70×10^5 disintegrations/s. The isotope has a half-life of 1.28×10^9 y.

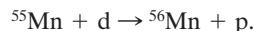
••35 SSM A certain radionuclide is being manufactured in a cyclotron at a constant rate R . It is also decaying with disintegration constant λ . Assume that the production process has been going on for a time that is much longer than the half-life of the radionuclide. (a) Show that the number of radioactive nuclei present after such time remains constant and is given by $N = R/\lambda$. (b) Now show that this result holds no matter how many radioactive nuclei were present initially. The nuclide is said to be in *secular equilibrium* with its source; in this state its decay rate is just equal to its production rate.

••36 Plutonium isotope ^{239}Pu decays by alpha decay with a half-life of 24 100 y. How many milligrams of helium are produced by an initially pure 12.0 g sample of ^{239}Pu at the end of 20 000 y? (Consider only the helium produced directly by the plutonium and not by any by-products of the decay process.)

••37 The radionuclide ^{64}Cu has a half-life of 12.7 h. If a sample contains 5.50 g of initially pure ^{64}Cu at $t = 0$, how much of it will decay between $t = 14.0$ h and $t = 16.0$ h?

••38 A dose of $8.60 \mu\text{Ci}$ of a radioactive isotope is injected into a patient. The isotope has a half-life of 3.0 h. How many of the isotope parents are injected?

••39 The radionuclide ^{56}Mn has a half-life of 2.58 h and is produced in a cyclotron by bombarding a manganese target with deuterons. The target contains only the stable manganese isotope ^{55}Mn , and the manganese–deuteron reaction that produces ^{56}Mn is



If the bombardment lasts much longer than the half-life of ^{56}Mn , the activity of the ^{56}Mn produced in the target reaches a final value of $8.88 \times 10^{10} \text{ Bq}$. (a) At what rate is ^{56}Mn being produced? (b) How many ^{56}Mn nuclei are then in the target? (c) What is their total mass?

••40 A source contains two phosphorus radionuclides, ^{32}P ($T_{1/2} = 14.3$ d) and ^{33}P ($T_{1/2} = 25.3$ d). Initially, 10.0% of the decays come from ^{33}P . How long must one wait until 90.0% do so?

••41 A 1.00 g sample of samarium emits alpha particles at a rate of 120 particles/s. The responsible isotope is ^{147}Sm , whose natural abundance in bulk samarium is 15.0%. Calculate the half-life for the decay process.

••42 What is the activity of a 20 ng sample of ^{92}Kr , which has a half-life of 1.84 s?

••43 A radioactive sample intended for irradiation of a hospital patient is prepared at a nearby laboratory. The sample has a half-life of 83.61 h. What should its initial activity be if its activity is to be $7.4 \times 10^8 \text{ Bq}$ when it is used to irradiate the patient 24 h later?

••44 GO Figure 42-19 shows the decay of parents in a radioactive sample. The axes are scaled by $N_s = 2.00 \times 10^6$ and $t_s = 10.0$ s. What is the activity of the sample at $t = 27.0$ s?

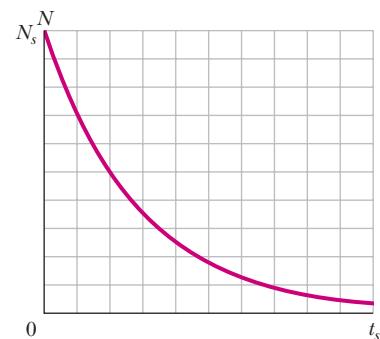


Fig. 42-19 Problem 44.

••45 In 1992, Swiss police arrested two men who were attempting to smuggle osmium out of Eastern Europe for a clandestine sale. However, by error, the smugglers had picked up ^{137}Cs . Reportedly, each smuggler was carrying a 1.0 g sample of ^{137}Cs in a pocket! In (a) bequerels and (b) curies, what was the activity of each sample? The isotope ^{137}Cs has a half-life of 30.2 y. (The activities of radioisotopes commonly used in hospitals range up to a few millicuries.)

••46 The radioactive nuclide ^{99}Tc can be injected into a patient's bloodstream in order to monitor the blood flow, measure the blood volume, or find a tumor, among other goals. The nuclide is produced in a hospital by a "cow" containing ^{99}Mo , a radioactive nuclide that decays to ^{99}Tc with a half-life of 67 h. Once a day, the cow is "milked" for its ^{99}Tc , which is produced in an excited state by the ^{99}Mo ; the ^{99}Tc de-excites to its lowest energy state by emitting a gamma-ray photon, which is recorded by detectors placed around the patient. The de-excitation has a half-life of 6.0 h. (a) By what process does ^{99}Mo decay to ^{99}Tc ? (b) If a patient is injected with an $8.2 \times 10^7 \text{ Bq}$ sample of ^{99}Tc , how many gamma-ray photons are initially produced within the patient each second? (c) If the emission rate of gamma-ray photons from a small tumor that has collected ^{99}Tc is 38 per second at a certain time, how many excited-state ^{99}Tc are located in the tumor at that time?

••47 SSM After long effort, in 1902 Marie and Pierre Curie succeeded in separating from uranium ore the first substantial quantity of radium, one decigram of pure RaCl_2 . The radium was the radioactive isotope ^{226}Ra , which has a half-life of 1600 y. (a) How many radium nuclei had the Curies isolated? (b) What was the decay rate of their sample, in disintegrations per second?

sec. 42-5 Alpha Decay

••48 How much energy is released when a ^{238}U nucleus decays by emitting (a) an alpha particle and (b) a sequence of neutron, proton, neutron, proton? (c) Convince yourself both by reasoned argument and by direct calculation that the difference between these two numbers is just the total binding energy of the alpha particle. (d) Find that binding energy. Some needed atomic and particle masses are

1192 CHAPTER 42 NUCLEAR PHYSICS

^{238}U	238.050 79 u	^{234}Th	234.043 63 u
^{237}U	237.048 73 u	^4He	4.002 60 u
^{236}Pa	236.048 91 u	^1H	1.007 83 u
^{235}Pa	235.045 44 u	n	1.008 66 u

•49 SSM Generally, more massive nuclides tend to be more unstable to alpha decay. For example, the most stable isotope of uranium, ^{238}U , has an alpha decay half-life of 4.5×10^9 y. The most stable isotope of plutonium is ^{244}Pu with an 8.0×10^7 y half-life, and for curium we have ^{248}Cm and 3.4×10^5 y. When half of an original sample of ^{238}U has decayed, what fraction of the original sample of (a) plutonium and (b) curium is left?

•50 Large radionuclides emit an alpha particle rather than other combinations of nucleons because the alpha particle has such a stable, tightly bound structure. To confirm this statement, calculate the disintegration energies for these hypothetical decay processes and discuss the meaning of your findings:

- (a) $^{235}\text{U} \rightarrow ^{232}\text{Th} + ^3\text{He}$, (b) $^{235}\text{U} \rightarrow ^{231}\text{Th} + ^4\text{He}$,
 (c) $^{235}\text{U} \rightarrow ^{230}\text{Th} + ^5\text{He}$.

The needed atomic masses are

^{232}Th	232.0381 u	^3He	3.0160 u
^{231}Th	231.0363 u	^4He	4.0026 u
^{230}Th	230.0331 u	^5He	5.0122 u
^{235}U	235.0429 u		

•51 A ^{238}U nucleus emits a 4.196 MeV alpha particle. Calculate the disintegration energy Q for this process, taking the recoil energy of the residual ^{234}Th nucleus into account.

•52 Under certain rare circumstances, a nucleus can decay by emitting a particle more massive than an alpha particle. Consider the decays



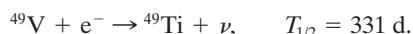
Calculate the Q value for the (a) first and (b) second decay and determine that both are energetically possible. (c) The Coulomb barrier height for alpha-particle emission is 30.0 MeV. What is the barrier height for ^{14}C emission? The needed atomic masses are

^{223}Ra	223.018 50 u	^{14}C	14.003 24 u
^{209}Pb	208.981 07 u	^4He	4.002 60 u
^{219}Rn	219.009 48 u		

sec. 42-6 Beta Decay

•53 SSM The cesium isotope ^{137}Cs is present in the fallout from aboveground detonations of nuclear bombs. Because it decays with a slow (30.2 y) half-life into ^{137}Ba , releasing considerable energy in the process, it is of environmental concern. The atomic masses of the Cs and Ba are 136.9071 and 136.9058 u, respectively; calculate the total energy released in such a decay.

•54 Some radionuclides decay by capturing one of their own atomic electrons, a K-shell electron, say. An example is



Show that the disintegration energy Q for this process is given by

$$Q = (m_V - m_{\text{Ti}})c^2 - E_K,$$

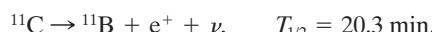
where m_V and m_{Ti} are the atomic masses of ^{49}V and ^{49}Ti , respec-

tively, and E_K is the binding energy of the vanadium K-shell electron. (Hint: Put \mathbf{m}_V and \mathbf{m}_{Ti} as the corresponding nuclear masses and then add in enough electrons to use the atomic masses.)

•55 A free neutron decays according to Eq. 42-26. If the neutron–hydrogen atom mass difference is $840 \mu\text{u}$, what is the maximum kinetic energy K_{max} possible for the electron produced in a neutron decay?

•56 An electron is emitted from a middle-mass nuclide ($A = 150$, say) with a kinetic energy of 1.0 MeV. (a) What is its de Broglie wavelength? (b) Calculate the radius of the emitting nucleus. (c) Can such an electron be confined as a standing wave in a “box” of such dimensions? (d) Can you use these numbers to disprove the (abandoned) argument that electrons actually exist in nuclei?

•57 GO The radionuclide ^{11}C decays according to



The maximum energy of the emitted positrons is 0.960 MeV. (a) Show that the disintegration energy Q for this process is given by

$$Q = (m_C - m_B - 2m_e)c^2,$$

where m_C and m_B are the atomic masses of ^{11}C and ^{11}B , respectively, and m_e is the mass of a positron. (b) Given the mass values $m_C = 11.011 424$ u, $m_B = 11.009 305$ u, and $m_e = 0.000 548 6$ u, calculate Q and compare it with the maximum energy of the emitted positron given above. (Hint: Let \mathbf{m}_C and \mathbf{m}_B be the nuclear masses and then add in enough electrons to use the atomic masses.)

•58 Two radioactive materials that alpha decay, ^{238}U and ^{232}Th , and one that beta decays, ^{40}K , are sufficiently abundant in granite to contribute significantly to the heating of Earth through the decay energy produced. The alpha-decay isotopes give rise to decay chains that stop when stable lead isotopes are formed. The isotope ^{40}K has a single beta decay. (Assume this is the only possible decay of that isotope.) Here is the information:

Parent	Decay Mode	Half-Life (y)	Stable End Point		
			Q (MeV)	f	(ppm)
^{238}U	α	4.47×10^9	^{206}Pb	51.7	4
^{232}Th	α	1.41×10^{10}	^{208}Pb	42.7	13
^{40}K	β	1.28×10^9	^{40}Ca	1.31	4

In the table Q is the *total* energy released in the decay of one parent nucleus to the *final* stable end point and f is the abundance of the isotope in kilograms per kilogram of granite; ppm means parts per million. (a) Show that these materials produce energy as heat at the rate of 1.0×10^{-9} W for each kilogram of granite. (b) Assuming that there is 2.7×10^{22} kg of granite in a 20-km-thick spherical shell at the surface of Earth, estimate the power of this decay process over all of Earth. Compare this power with the total solar power intercepted by Earth, 1.7×10^{17} W.

•59 SSM WWW The radionuclide ^{32}P decays to ^{32}S as described by Eq. 42-24. In a particular decay event, a 1.71 MeV electron is emitted, the maximum possible value. What is the kinetic energy of the recoiling ^{32}S atom in this event? (Hint: For the electron it is necessary to use the relativistic expressions for kinetic energy and linear momentum. The ^{32}S atom is nonrelativistic.)

sec. 42-7 Radioactive Dating

•60 A 5.00 g charcoal sample from an ancient fire pit has a ^{14}C activity of 63.0 disintegrations/min. A living tree has a ^{14}C activity of 15.3 disintegrations/min per 1.00 g. The half-life of ^{14}C is 5730 y. How old is the charcoal sample?

•61 The isotope ^{238}U decays to ^{206}Pb with a half-life of 4.47×10^9 y. Although the decay occurs in many individual steps, the first step has by far the longest half-life; therefore, one can often consider the decay to go directly to lead. That is,



A rock is found to contain 4.20 mg of ^{238}U and 2.135 mg of ^{206}Pb . Assume that the rock contained no lead at formation, so all the lead now present arose from the decay of uranium. How many atoms of (a) ^{238}U and (b) ^{206}Pb does the rock now contain? (c) How many atoms of ^{238}U did the rock contain at formation? (d) What is the age of the rock?

••62 A particular rock is thought to be 260 million years old. If it contains 3.70 mg of ^{238}U , how much ^{206}Pb should it contain? See Problem 61.

••63 A rock recovered from far underground is found to contain 0.86 mg of ^{238}U , 0.15 mg of ^{206}Pb , and 1.6 mg of ^{40}Ar . How much ^{40}K will it likely contain? Assume that ^{40}K decays to only ^{40}Ar with a half-life of 1.25×10^9 y. Also assume that ^{238}U has a half-life of 4.47×10^9 y.

••64 The isotope ^{40}K can decay to either ^{40}Ca or ^{40}Ar ; assume both decays have a half-life of 1.26×10^9 y. The ratio of the Ca produced to the Ar produced is $8.54/1 = 8.54$. A sample originally had only ^{40}K . It now has equal amounts of ^{40}K and ^{40}Ar ; that is, the ratio of K to Ar is $1/1 = 1$. How old is the sample? (*Hint:* Work this like other radioactive-dating problems, except that this decay has two products.)

sec. 42-8 Measuring Radiation Dosage

•65 The nuclide ^{198}Au , with a half-life of 2.70 d, is used in cancer therapy. What mass of this nuclide is required to produce an activity of 250 Ci?

•66 A radiation detector records 8700 counts in 1.00 min. Assuming that the detector records all decays, what is the activity of the radiation source in (a) becquerels and (b) curies?

•67 An organic sample of mass 4.00 kg absorbs 2.00 mJ via slow neutron radiation (RBE = 5). What is the dose equivalent (mSv)?

•68 A 75 kg person receives a whole-body radiation dose of 2.4×10^{-4} Gy, delivered by alpha particles for which the RBE factor is 12. Calculate (a) the absorbed energy in joules and the dose equivalent in (b) sieverts and (c) rem.

•69 An 85 kg worker at a breeder reactor plant accidentally ingests 2.5 mg of ^{239}Pu dust. This isotope has a half-life of 24 100 y, decaying by alpha decay. The energy of the emitted alpha particles is 5.2 MeV, with an RBE factor of 13. Assume that the plutonium resides in the worker's body for 12 h (it is eliminated naturally by the digestive system rather than being absorbed by any of the internal organs) and that 95% of the emitted alpha particles are stopped within the body. Calculate (a) the number of plutonium atoms ingested, (b) the number that decay during the 12 h, (c) the energy absorbed by the body, (d) the resulting physical dose in grays, and (e) the dose equivalent in sieverts.

sec. 42-9 Nuclear Models

•70 A typical kinetic energy for a nucleon in a middle-mass nucleus may be taken as 5.00 MeV. To what effective nuclear temperature does this correspond, based on the assumptions of the collective model of nuclear structure?

•71 A measurement of the energy E of an intermediate nucleus must be made within the mean lifetime Δt of the nucleus and necessarily carries an uncertainty ΔE according to the uncertainty principle

$$\Delta E \cdot \Delta t = \hbar$$

(a) What is the uncertainty ΔE in the energy for an intermediate nucleus if the nucleus has a mean lifetime of 10^{-22} s? (b) Is the nucleus a compound nucleus?

•72 In the following list of nuclides, identify (a) those with filled nucleon shells, (b) those with one nucleon outside a filled shell, and (c) those with one vacancy in an otherwise filled shell: ^{13}C , ^{18}O , ^{40}K , ^{49}Ti , ^{60}Ni , ^{91}Zr , ^{92}Mo , ^{121}Sb , ^{143}Nd , ^{144}Sm , ^{205}Tl , and ^{207}Pb .

••73 Consider the three formation processes shown for the compound nucleus ^{20}Ne in Fig. 42-14. Here are some of the atomic and particle masses:

^{20}Ne	19.992 44 u	α	4.002 60 u
^{19}F	18.998 40 u	p	1.007 83 u
^{16}O	15.994 91 u		

What energy must (a) the alpha particle, (b) the proton, and (c) the γ -ray photon have to provide 25.0 MeV of excitation energy to the compound nucleus?

Additional Problems

74 In a certain rock, the ratio of lead atoms to uranium atoms is 0.300. Assume that uranium has a half-life of 4.47×10^9 y and that the rock had no lead atoms when it formed. How old is the rock?

75 A certain stable nuclide, after absorbing a neutron, emits an electron, and the new nuclide splits spontaneously into two alpha particles. Identify the nuclide.

76 A typical chest x-ray radiation dose is 250 μSv , delivered by x rays with an RBE factor of 0.85. Assuming that the mass of the exposed tissue is one-half the patient's mass of 88 kg, calculate the energy absorbed in joules.

77 How many years are needed to reduce the activity of ^{14}C to 0.020 of its original activity? The half-life of ^{14}C is 5730 y.

78 Radioactive element AA can decay to either element BB or element CC . The decay depends on chance, but the ratio of the resulting number of BB atoms to the resulting number of CC atoms is always 2/1. The decay has a half-life of 8.00 days. We start with a sample of pure AA . How long must we wait until the number of CC atoms is 1.50 times the number of AA atoms?

79 One of the dangers of radioactive fallout from a nuclear bomb is its ^{90}Sr , which decays with a 29-year half-life. Because it has chemical properties much like those of calcium, the strontium, if ingested by a cow, becomes concentrated in the cow's milk. Some of the ^{90}Sr ends up in the bones of whoever drinks the milk. The energetic electrons emitted in the beta decay of ^{90}Sr damage the bone marrow and thus impair the production of red blood cells. A 1 megaton bomb produces approximately 400 g of ^{90}Sr . If the fallout spreads uniformly over a 2000 km² area, what ground area would hold an amount of radioactivity equal to the "allowed" limit for one person, which is 74 000 counts/s?

1194 CHAPTER 42 NUCLEAR PHYSICS

80 Because of the 1986 explosion and fire in a reactor at the Chernobyl nuclear power plant in northern Ukraine, part of Ukraine is contaminated with ^{137}Cs , which undergoes beta-minus decay with a half-life of 30.2 y. In 1996, the total activity of this contamination over an area of $2.6 \times 10^5 \text{ km}^2$ was estimated to be $1 \times 10^{16} \text{ Bq}$. Assume that the ^{137}Cs is uniformly spread over that area and that the beta-decay electrons travel either directly upward or directly downward. How many beta-decay electrons would you intercept were you to lie on the ground in that area for 1 h (a) in 1996 and (b) today? (You need to estimate your cross-sectional area that intercepts those electrons.)

81 Figure 42-20 shows part of the decay scheme of ^{237}Np on a plot of mass number A versus proton number Z ; five lines that represent either alpha decay or beta-minus decay connect dots that represent isotopes. What is the isotope at the end of the five decays (as marked with a question mark in Fig. 42-20)?

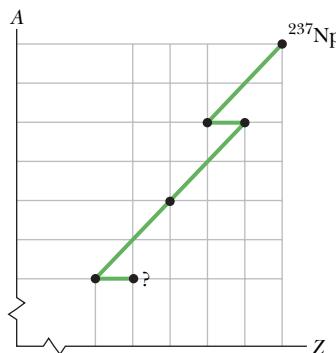


Fig. 42-20 Problem 81.

82 After a brief neutron irradiation of silver, two isotopes are present: ^{108}Ag ($T_{1/2} = 2.42 \text{ min}$) with an initial decay rate of $3.1 \times 10^5/\text{s}$, and ^{110}Ag ($T_{1/2} = 24.6 \text{ s}$) with an initial decay rate of $4.1 \times 10^6/\text{s}$. Make a semilog plot similar to Fig. 42-9 showing the total combined decay rate of the two isotopes as a function of time from $t = 0$ until $t = 10 \text{ min}$. We used Fig. 42-9 to illustrate the extraction of the half-life for simple (one isotope) decays. Given only your plot of total decay rate for the two-isotope system here, suggest a way to analyze it in order to find the half-lives of both isotopes.

83 Because a nucleon is confined to a nucleus, we can take the uncertainty in its position to be approximately the nuclear radius r . Use the uncertainty principle to determine the uncertainty Δp in the linear momentum of the nucleon. Using the approximation $p \approx \Delta p$ and the fact that the nucleon is nonrelativistic, calculate the kinetic energy of the nucleon in a nucleus with $A = 100$.

84 A radium source contains 1.00 mg of ^{226}Ra , which decays with a half-life of 1600 y to produce ^{222}Rn , a noble gas. This radon isotope in turn decays by alpha emission with a half-life of 3.82 d. If this process continues for a time much longer than the half-life of ^{222}Rn , the ^{222}Rn decay rate reaches a limiting value that matches the rate at which ^{222}Rn is being produced, which is approximately constant because of the relatively long half-life of ^{226}Ra . For the source under this limiting condition, what are (a) the activity of ^{226}Ra , (b) the activity of ^{222}Rn , and (c) the total mass of ^{222}Rn ?

85 Make a nuclidic chart similar to Fig. 42-6 for the 25 nuclides $^{118-122}\text{Te}$, $^{117-121}\text{Sb}$, $^{116-120}\text{Sn}$, $^{115-119}\text{In}$, and $^{114-118}\text{Cd}$. Draw in and label (a) all isobaric (constant A) lines and (b) all lines of constant neutron excess, defined as $N - Z$.

86 GO A projectile alpha particle is headed directly toward a target aluminum nucleus. Both objects are assumed to be spheres. What energy is required of the alpha particle if it is to momentarily stop just as its “surface” touches the “surface” of the aluminum nucleus? Assume that the target nucleus remains stationary.

87 Consider a ^{238}U nucleus to be made up of an alpha particle (^4He) and a residual nucleus (^{234}Th). Plot the electrostatic potential energy $U(r)$, where r is the distance between these particles. Cover the approximate range $10 \text{ fm} < r < 100 \text{ fm}$ and compare your plot with that of Fig. 42-10.

88 Characteristic nuclear time is a useful but loosely defined quantity, taken to be the time required for a nucleon with a few million electron-volts of kinetic energy to travel a distance equal to the diameter of a middle-mass nuclide. What is the order of magnitude of this quantity? Consider 5 MeV neutrons traversing a nuclear diameter of ^{197}Au ; use Eq. 42-3.

89 The radius of a spherical nucleus is measured, by electron-scattering methods, to be 3.6 fm. What is the likely mass number of the nucleus?

90 Using a nuclidic chart, write the symbols for (a) all stable isotopes with $Z = 60$, (b) all radioactive nuclides with $N = 60$, and (c) all nuclides with $A = 60$.

91 If the unit for atomic mass were defined so that the mass of ^1H were exactly 1.000 000 u, what would be the mass of (a) ^{12}C (actual mass 12.000 000 u) and (b) ^{238}U (actual mass 238.050 785 u)?

92 High-mass radionuclides, which may be either alpha or beta emitters, belong to one of four decay chains, depending on whether their mass number A is of the form $4n$, $4n + 1$, $4n + 2$, or $4n + 3$, where n is a positive integer. (a) Justify this statement and show that if a nuclide belongs to one of these families, all its decay products belong to the same family. Classify the following nuclides as to family: (b) ^{235}U , (c) ^{236}U , (d) ^{238}U , (e) ^{239}Pu , (f) ^{240}Pu , (g) ^{245}Cm , (h) ^{246}Cm , (i) ^{249}Cf , and (j) ^{253}Fm .

93 Find the disintegration energy Q for the decay of ^{49}V by K -electron capture (see Problem 54). The needed data are $m_V = 48.948\ 52 \text{ u}$, $m_{\text{Ti}} = 48.947\ 87 \text{ u}$, and $E_K = 5.47 \text{ keV}$.

94 Locate the nuclides displayed in Table 42-1 on the nuclidic chart of Fig. 42-5. Verify that they lie in the stability zone.

95 The radionuclide ^{32}P ($T_{1/2} = 14.28 \text{ d}$) is often used as a tracer to follow the course of biochemical reactions involving phosphorus. (a) If the counting rate in a particular experimental setup is initially 3050 counts/s, how much time will the rate take to fall to 170 counts/s? (b) A solution containing ^{32}P is fed to the root system of an experimental tomato plant, and the ^{32}P activity in a leaf is measured 3.48 days later. By what factor must this reading be multiplied to correct for the decay that has occurred since the experiment began?

96 From data presented in the first few paragraphs of Section 42-4, find (a) the disintegration constant λ and (b) the half-life of ^{238}U .

ENERGY FROM THE NUCLEUS

43-1 WHAT IS PHYSICS?

Let's now turn to a central concern of physics and certain types of engineering: Can we get useful energy from nuclear sources, as people have done for thousands of years from atomic sources by burning materials like wood and coal? As you already know, the answer is yes, but there are major differences between the two energy sources. When we get energy from wood and coal by burning them, we are tinkering with atoms of carbon and oxygen, rearranging their outer *electrons* into more stable combinations. When we get energy from uranium in a nuclear reactor, we are again burning a fuel, but now we are tinkering with the uranium nucleus, rearranging its *nucleons* into more stable combinations.

Electrons are held in atoms by the electromagnetic Coulomb force, and it takes only a few electron-volts to pull one of them out. On the other hand, nucleons are held in nuclei by the strong force, and it takes a few *million* electron-volts to pull one of *them* out. This factor of a few million is reflected in the fact that we can extract a few million times more energy from a kilogram of uranium than we can from a kilogram of coal.

In both atomic and nuclear burning, the release of energy is accompanied by a decrease in mass, according to the equation $Q = -\Delta m c^2$. The central difference between burning uranium and burning coal is that, in the former case, a much larger fraction of the available mass (again, by a factor of a few million) is consumed.

The different processes that can be used for atomic or nuclear burning provide different levels of power, or rates at which the energy is delivered. In the nuclear case, we can burn a kilogram of uranium explosively in a bomb or slowly in a power reactor. In the atomic case, we might consider exploding a stick of dynamite or digesting a jelly doughnut.

Table 43-1 shows how much energy can be extracted from 1 kg of matter by doing various things to it. Instead of reporting the energy directly, the table

Table 43-1

Energy Released by 1 kg of Matter

Form of Matter	Process	Time ^a
Water	A 50 m waterfall	5 s
Coal	Burning	8 h
Enriched UO_2	Fission in a reactor	690 y
^{235}U	Complete fission	3×10^4 y
Hot deuterium gas	Complete fusion	3×10^4 y
Matter and antimatter	Complete annihilation	3×10^7 y

^aThis column shows the time interval for which the generated energy could power a 100 W lightbulb.

shows how long the extracted energy could operate a 100 W lightbulb. Only processes in the first three rows of the table have actually been carried out; the remaining three represent theoretical limits that may not be attainable in practice. The bottom row, the total mutual annihilation of matter and antimatter, is an ultimate energy production goal. In that process, *all* the mass energy is transferred to other forms of energy.

The comparisons of Table 43-1 are computed on a per-unit-mass basis. Kilogram for kilogram, you get several million times more energy from uranium than you do from coal or from falling water. On the other hand, there is a lot of coal in Earth's crust, and water is easily backed up behind a dam.

43-2 Nuclear Fission: The Basic Process

In 1932 English physicist James Chadwick discovered the neutron. A few years later Enrico Fermi in Rome found that when various elements are bombarded by neutrons, new radioactive elements are produced. Fermi had predicted that the neutron, being uncharged, would be a useful nuclear projectile; unlike the proton or the alpha particle, it experiences no repulsive Coulomb force when it nears a nuclear surface. Even *thermal neutrons*, which are slowly moving neutrons in thermal equilibrium with the surrounding matter at room temperature, with a kinetic energy of only about 0.04 eV, are useful projectiles in nuclear studies.

In the late 1930s physicist Lise Meitner and chemists Otto Hahn and Fritz Strassmann, working in Berlin and following up on the work of Fermi and his co-workers, bombarded solutions of uranium salts with such thermal neutrons. They found that after the bombardment a number of new radionuclides were present. In 1939 one of the radionuclides produced in this way was positively identified, by repeated tests, as barium. But how, Hahn and Strassmann wondered, could this middle-mass element ($Z = 56$) be produced by bombarding uranium ($Z = 92$) with neutrons?

The puzzle was solved within a few weeks by Meitner and her nephew Otto Frisch. They suggested the mechanism by which a uranium nucleus, having absorbed a thermal neutron, could split, with the release of energy, into two roughly equal parts, one of which might well be barium. Frisch named the process **fission**.

Meitner's central role in the discovery of fission was not fully recognized until recent historical research brought it to light. She did not share in the Nobel Prize in chemistry that was awarded to Otto Hahn in 1944. However, Meitner was honored by having an element named after her: meitnerium (symbol Mt, $Z = 109$).

A Closer Look at Fission

Figure 43-1 shows the distribution by mass number of the fragments produced when ^{235}U is bombarded with thermal neutrons. The most probable mass numbers, occurring in about 7% of the events, are centered around $A \approx 95$ and $A \approx 140$. Curiously, the “double-peaked” character of Fig. 43-1 is still not understood.

In a typical ^{235}U fission event, a ^{235}U nucleus absorbs a thermal neutron, producing a compound nucleus ^{236}U in a highly excited state. It is *this* nucleus that actually undergoes fission, splitting into two fragments. These fragments—between them—rapidly emit two neutrons, leaving (in a typical case) ^{140}Xe ($Z = 54$) and ^{94}Sr ($Z = 38$) as fission fragments. Thus, the stepwise fission equation for this event is



Note that during the formation and fission of the compound nucleus, there is

43-2 NUCLEAR FISSION: THE BASIC PROCESS

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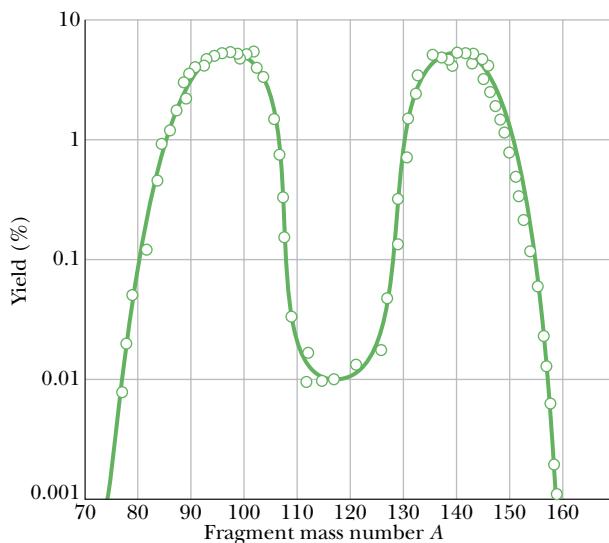


Fig. 43-1 The distribution by mass number of the fragments that are found when many fission events of ^{235}U are examined. Note that the vertical scale is logarithmic.

conservation of the number of protons and of the number of neutrons involved in the process (and thus conservation of their total number and the net charge).

In Eq. 43-1, the fragments ^{140}Xe and ^{94}Sr are both highly unstable, undergoing beta decay (with the conversion of a neutron to a proton and the emission of an electron and a neutrino) until each reaches a stable end product. For xenon, the decay chain is

^{140}Xe	\rightarrow	^{140}Cs	\rightarrow	^{140}Ba	\rightarrow	^{140}La	\rightarrow	^{140}Ce	
$T_{1/2}$	14 s	64 s	13 d	40 h		Stable			(43-2)
Z	54	55	56	57		58			

For strontium, it is

^{94}Sr	\rightarrow	^{94}Y	\rightarrow	^{94}Zr	
$T_{1/2}$	75 s	19 min		Stable	(43-3)
Z	38	39		40	

As we should expect from Section 42-6, the mass numbers (140 and 94) of the fragments remain unchanged during these beta-decay processes and the atomic numbers (initially 54 and 38) increase by unity at each step.

Inspection of the stability band on the nuclidic chart of Fig. 42-5 shows why the fission fragments are unstable. The nuclide ^{236}U , which is the fissioning nucleus in the reaction of Eq. 43-1, has 92 protons and $236 - 92$, or 144, neutrons, for a neutron/proton ratio of about 1.6. The primary fragments formed immediately after the fission reaction have about this same neutron/proton ratio. However, stable nuclides in the middle-mass region have smaller neutron/proton ratios, in the range of 1.3 to 1.4. The primary fragments are thus *neutron rich* (they have too many neutrons) and will eject a few neutrons, two in the case of the reaction of Eq. 43-1. The fragments that remain are still too neutron rich to be stable. Beta decay offers a mechanism for getting rid of the excess neutrons—namely, by changing them into protons within the nucleus.

We can estimate the energy released by the fission of a high-mass nuclide by examining the total binding energy per nucleon ΔE_{ben} before and after the fission. The idea is that fission can occur because the total mass energy will

decrease; that is, ΔE_{ben} will *increase* so that the products of the fission are *more* tightly bound. Thus, the energy Q released by the fission is

$$Q = \left(\frac{\text{total final binding energy}}{\text{binding energy}} \right) - \left(\frac{\text{initial binding energy}}{\text{binding energy}} \right). \quad (43-4)$$

For our estimate, let us assume that fission transforms an initial high-mass nucleus to two middle-mass nuclei with the same number of nucleons. Then we have

$$Q = \left(\frac{\Delta E_{\text{ben}}}{\text{of nucleons}} \right) \left(\frac{\text{final number of nucleons}}{\text{initial number of nucleons}} \right). \quad (43-5)$$

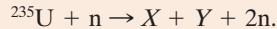
From Fig. 42-7, we see that for a high-mass nuclide ($A \approx 240$), the binding energy per nucleon is about 7.6 MeV/nucleon. For middle-mass nuclides ($A \approx 120$), it is about 8.5 MeV/nucleon. Thus, the energy released by fission of a high-mass nuclide to two middle-mass nuclides is

$$Q = \left(8.5 \frac{\text{MeV}}{\text{nucleon}} \right) (2 \text{ nuclei}) \left(120 \frac{\text{nucleons}}{\text{nucleus}} \right) - \left(7.6 \frac{\text{MeV}}{\text{nucleon}} \right) (240 \text{ nucleons}) \approx 200 \text{ MeV}. \quad (43-6)$$



CHECKPOINT 1

A generic fission event is



Which of the following pairs *cannot* represent X and Y : (a) ^{141}Xe and ^{93}Sr ; (b) ^{139}Cs and ^{95}Rb ; (c) ^{156}Nd and ^{79}Ge ; (d) ^{121}In and ^{113}Ru ?

Sample Problem

Q value in a fission of uranium-235

Find the disintegration energy Q for the fission event of Eq. 43-1, taking into account the decay of the fission fragments as displayed in Eqs. 43-2 and 43-3. Some needed atomic and particle masses are

^{235}U	235.0439 u	^{140}Ce	139.9054 u
n	1.008 66 u	^{94}Zr	93.9063 u

KEY IDEAS

- (1) The disintegration energy Q is the energy transferred from mass energy to kinetic energy of the decay products.
- (2) $Q = -\Delta m c^2$, where Δm is the change in mass.

Calculations: Because we are to include the decay of the fission fragments, we combine Eqs. 43-1, 43-2, and 43-3 to write the overall transformation as



Only the single neutron appears here because the initiating neutron on the left side of Eq. 43-1 cancels one of the two

neutrons on the right of that equation. The mass difference for the reaction of Eq. 43-7 is

$$\begin{aligned} \Delta m &= (139.9054 \text{ u} + 93.9063 \text{ u} + 1.008 66 \text{ u}) \\ &\quad - (235.0439 \text{ u}) \\ &= -0.223 54 \text{ u}, \end{aligned}$$

and the corresponding disintegration energy is

$$\begin{aligned} Q &= -\Delta m c^2 = -(-0.223 54 \text{ u})(931.494 013 \text{ MeV/u}) \\ &= 208 \text{ MeV}, \end{aligned} \quad (\text{Answer})$$

which is in good agreement with our estimate of Eq. 43-6.

If the fission event takes place in a bulk solid, most of this disintegration energy, which first goes into kinetic energy of the decay products, appears eventually as an increase in the internal energy of that body, revealing itself as a rise in temperature. Five or six percent or so of the disintegration energy, however, is associated with neutrinos that are emitted during the beta decay of the primary fission fragments. This energy is carried out of the system and is lost.

43-3 A Model for Nuclear Fission

Soon after the discovery of fission, Niels Bohr and John Wheeler used the collective model of the nucleus (Section 42-9), based on the analogy between a nucleus and a charged liquid drop, to explain the main nuclear features. Figure 43-2 suggests how the fission process proceeds from this point of view. When a high-mass nucleus—let us say ^{235}U —absorbs a slow (thermal) neutron, as in Fig. 43-2a, that neutron falls into the potential well associated with the strong forces that act in the nuclear interior. The neutron's potential energy is then transformed into internal excitation energy of the nucleus, as Fig. 43-2b suggests. The amount of excitation energy that a slow neutron carries into a nucleus is equal to the binding energy E_n of the neutron in that nucleus, which is the change in mass energy of the neutron–nucleus system due to the neutron's capture.

Figures 43-2c and d show that the nucleus, behaving like an energetically oscillating charged liquid drop, will sooner or later develop a short “neck” and will begin to separate into two charged “globs.” If the electric repulsion between these two globs forces them far enough apart to break the neck, the two fragments, each still carrying some residual excitation energy, will fly apart (Figs. 43-2e and f). Fission has occurred.

This model gave a good qualitative picture of the fission process. What remained to be seen, however, was whether it could answer a hard question: Why are some high-mass nuclides (^{235}U and ^{239}Pu , say) readily fissionable by thermal neutrons when other, equally massive nuclides (^{238}U and ^{243}Am , say) are not?

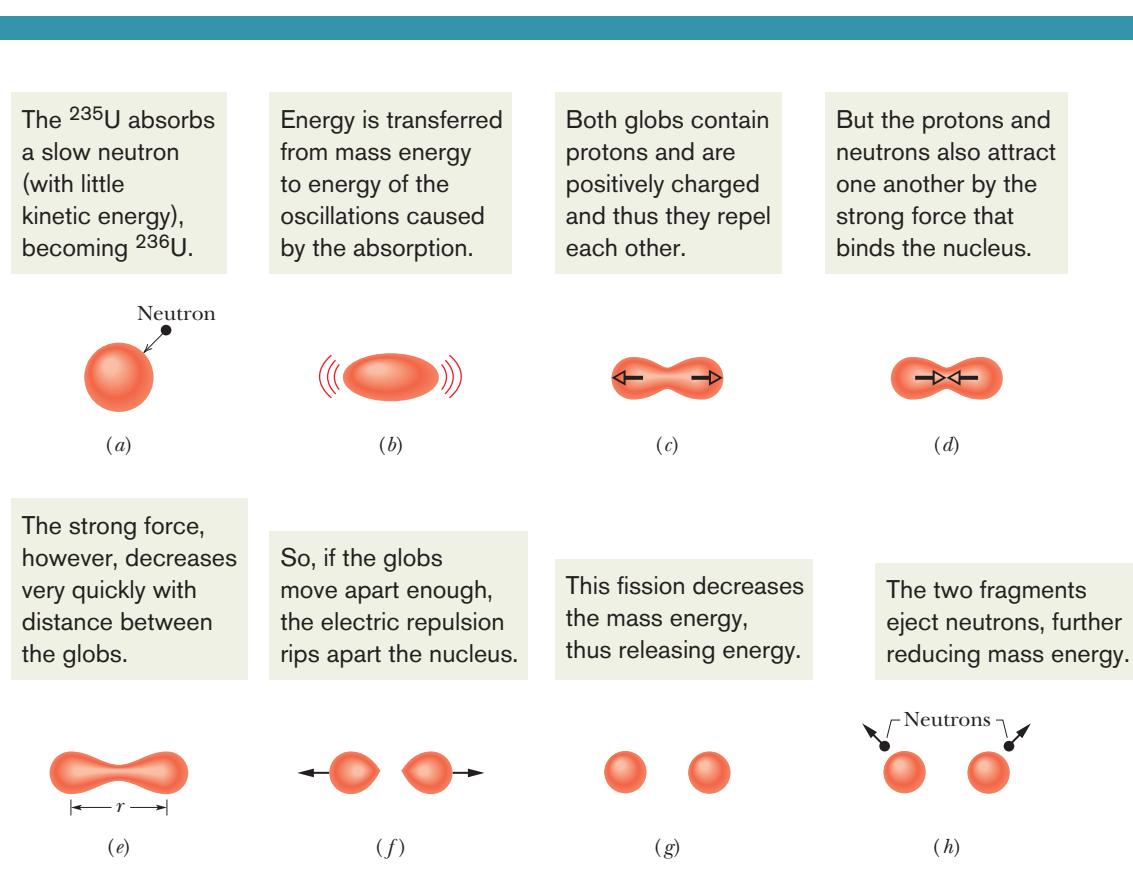


Fig. 43-2 The stages of a typical fission process, according to the collective model of Bohr and Wheeler.

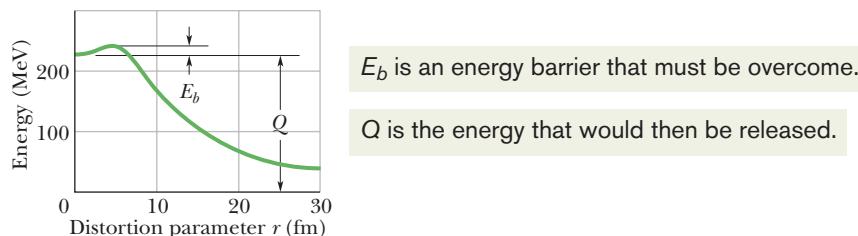


Fig. 43-3 The potential energy at various stages in the fission process, as predicted from the collective model of Bohr and Wheeler. The Q of the reaction (about 200 MeV) and the fission barrier height E_b are both indicated.

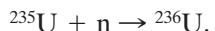
Bohr and Wheeler were able to answer this question. Figure 43-3 shows a graph of the potential energy of the fissioning nucleus at various stages, derived from their model for the fission process. This energy is plotted against the *distortion parameter* r , which is a rough measure of the extent to which the oscillating nucleus departs from a spherical shape. When the fragments are far apart, this parameter is simply the distance between their centers (Fig. 43-2e).

The energy difference between the initial state ($r = 0$) and the final state ($r = \infty$) of the fissioning nucleus—that is, the disintegration energy Q —is labeled in Fig. 43-3. The central feature of that figure, however, is that the potential energy curve passes through a maximum at a certain value of r . Thus, there is a *potential barrier* of height E_b that must be surmounted (or tunneled through) before fission can occur. This reminds us of alpha decay (Fig. 42-10), which is also a process that is inhibited by a potential barrier.

We see then that fission will occur only if the absorbed neutron provides an excitation energy E_n great enough to overcome the barrier. This energy E_n need not be *quite* as great as the barrier height E_b because of the possibility of quantum-physics tunneling.

Table 43-2 shows, for four high-mass nuclides, this test of whether capture of a thermal neutron can cause fissioning. For each nuclide, the table shows both the barrier height E_b of the nucleus that is formed by the neutron capture and the excitation energy E_n due to the capture. The values of E_b are calculated from the theory of Bohr and Wheeler. The values of E_n are calculated from the change in mass energy due to the neutron capture.

For an example of the calculation of E_n , we can go to the first line in the table, which represents the neutron capture process

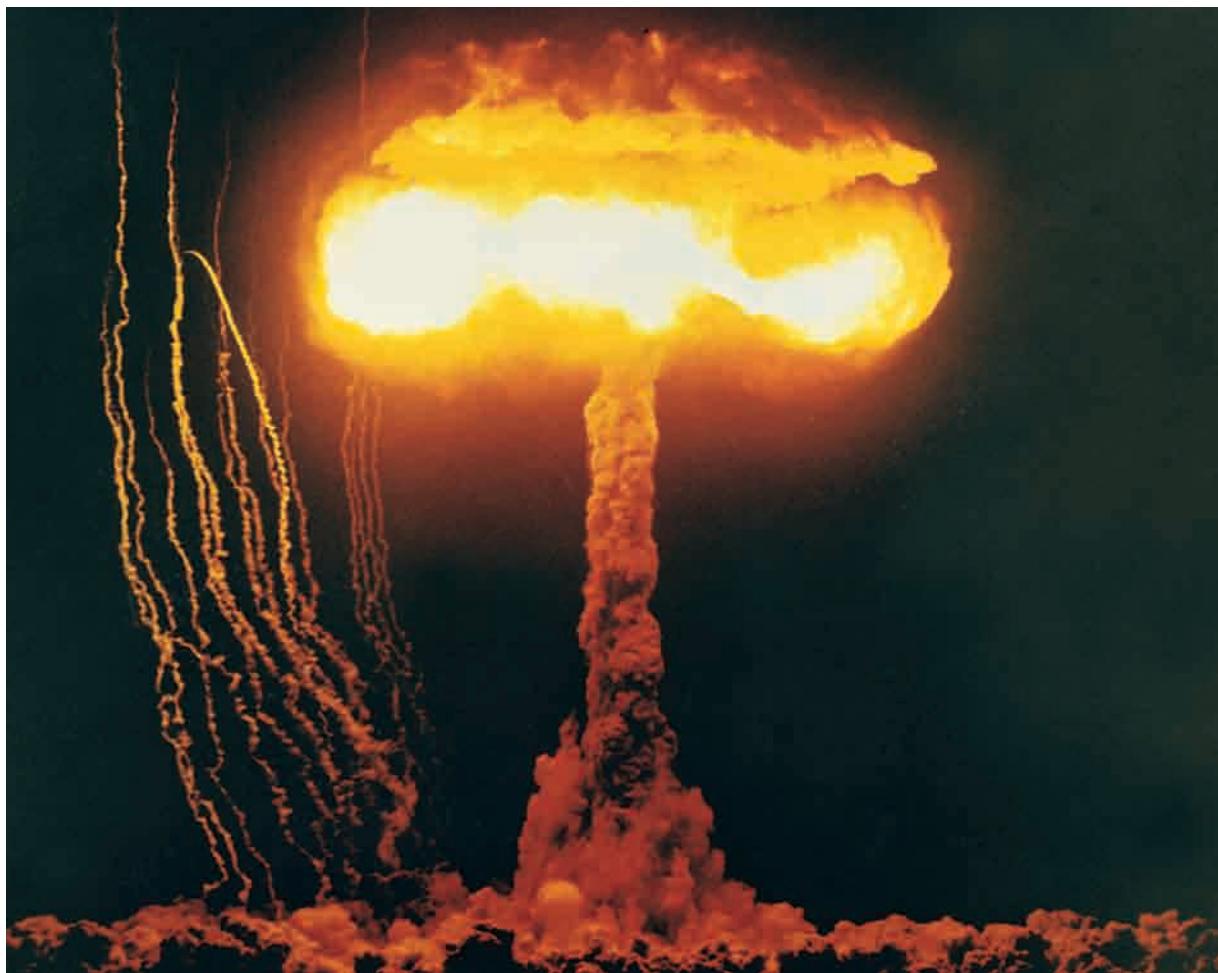


The masses involved are 235.043 922 u for ^{235}U , 1.008 665 u for the neutron, and 236.045 562 u for ^{236}U . It is easy to show that, because of the neutron capture, the mass decreases by 7.025×10^{-3} u. Thus, energy is transferred from mass energy to excitation energy E_n . Multiplying the change in mass by c^2 ($= 931.494\ 013$ MeV/u) gives us $E_n = 6.5$ MeV, which is listed on the first line of the table.

Table 43-2

Test of the Fissionability of Four Nuclides

Target Nuclide	Nuclide Being Fissioned	E_n (MeV)	E_b (MeV)	Fission by Thermal Neutrons?
^{235}U	^{236}U	6.5	5.2	Yes
^{238}U	^{239}U	4.8	5.7	No
^{239}Pu	^{240}Pu	6.4	4.8	Yes
^{243}Am	^{244}Am	5.5	5.8	No



Courtesy U.S. Department of Energy

Fig. 43-4 This image has transfixed the world since World War II. When Robert Oppenheimer, the head of the scientific team that developed the atomic bomb, witnessed the first atomic explosion, he quoted from a sacred Hindu text: “Now I am become Death, the destroyer of worlds.”

The first and third results in Table 43-2 are historically profound because they are the reasons the two atomic bombs used in World War II contained ^{235}U (first bomb) and ^{239}Pu (second bomb). That is, for ^{235}U and ^{239}Pu , $E_n > E_b$. This means that fission by absorption of a thermal neutron is predicted to occur for these nuclides. For the other two nuclides in Table 43-2 (^{238}U and ^{243}Am), we have $E_n < E_b$; thus, there is not enough energy from a thermal neutron for the excited nucleus to surmount the barrier or to tunnel through it effectively. Instead of fissioning, the nucleus gets rid of its excitation energy by emitting a gamma-ray photon.

The nuclides ^{238}U and ^{243}Am *can* be made to fission, however, if they absorb a substantially energetic (rather than a thermal) neutron. A ^{238}U nucleus, for example, might fission if it happens to absorb a neutron of at least 1.3 MeV in a so-called *fast fission* process (“fast” because the neutron is fast).

The two atomic bombs used in World War II depended on the ability of thermal neutrons to cause many high-mass nuclides in the cores of the bombs to fission nearly all at once, so that the fissioning would result in an explosive and devastating output of energy. Researchers knew that ^{235}U would work, but they had refined only enough for one bomb from uranium ore, which consists mainly of ^{238}U , which thermal neutrons will not fission. As the first bomb was being deployed, a ^{239}Pu bomb was tested successfully in New Mexico (Fig. 43-4), so the next deployed bomb contained ^{239}Pu rather than ^{235}U .

43-4 The Nuclear Reactor

For large-scale energy release due to fission, one fission event must trigger others, so that the process spreads throughout the nuclear fuel like flame through a log. The fact that more neutrons are produced in fission than are consumed raises the possibility of just such a **chain reaction**, with each neutron that is produced potentially triggering another fission. The reaction can be either rapid (as in a nuclear bomb) or controlled (as in a nuclear reactor).

Suppose that we wish to design a reactor based on the fission of ^{235}U by thermal neutrons. Natural uranium contains 0.7% of this isotope, the remaining 99.3% being ^{238}U , which is not fissionable by thermal neutrons. Let us give ourselves an edge by artificially *enriching* the uranium fuel so that it contains perhaps 3% ^{235}U . Three difficulties still stand in the way of a working reactor.

- 1. The Neutron Leakage Problem.** Some of the neutrons produced by fission will leak out of the reactor and so not be part of the chain reaction. Leakage is a surface effect; its magnitude is proportional to the square of a typical reactor dimension (the surface area of a cube of edge length a is $6a^2$). Neutron production, however, occurs throughout the volume of the fuel and is thus proportional to the cube of a typical dimension (the volume of the same cube is a^3). We can make the fraction of neutrons lost by leakage as small as we wish by making the reactor core large enough, thereby reducing the surface-to-volume ratio ($= 6/a$ for a cube).
- 2. The Neutron Energy Problem.** The neutrons produced by fission are fast, with kinetic energies of about 2 MeV. However, fission is induced most effectively by thermal neutrons. The fast neutrons can be slowed down by mixing the uranium fuel with a substance—called a **moderator**—that has two properties: It is effective in slowing down neutrons via elastic collisions, and it does not remove neutrons from the core by absorbing them so that they do not result in fission. Most power reactors in North America use water as a moderator; the hydrogen nuclei (protons) in the water are the effective component. We saw in Chapter 9 that if a moving particle has a head-on elastic collision with a stationary particle, the moving particle loses *all* its kinetic energy if the two particles have the same mass. Thus, protons form an effective moderator because they have approximately the same mass as the fast neutrons whose speed we wish to reduce.
- 3. The Neutron Capture Problem.** As the fast (2 MeV) neutrons generated by fission are slowed down in the moderator to thermal energies (about 0.04 eV), they must pass through a critical energy interval (from 1 to 100 eV) in which they are particularly susceptible to nonfission capture by ^{238}U nuclei. Such *resonance capture*, which results in the emission of a gamma ray, removes the neutron from the fission chain. To minimize such nonfission capture, the uranium fuel and the moderator are not intimately mixed but rather are placed in different regions of the reactor volume.

In a typical reactor, the uranium fuel is in the form of uranium oxide pellets, which are inserted end to end into long, hollow metal tubes. The liquid moderator surrounds bundles of these **fuel rods**, forming the reactor **core**. This geometric arrangement increases the probability that a fast neutron, produced in a fuel rod, will find itself in the moderator when it passes through the critical energy interval. Once the neutron has reached thermal energies, it may *still* be captured in ways that do not result in fission (called *thermal capture*). However, it is much more likely that the thermal neutron will wander back into a fuel rod and produce a fission event.

Figure 43-5 shows the neutron balance in a typical power reactor operating at constant power. Let us trace a sample of 1000 thermal neutrons through one

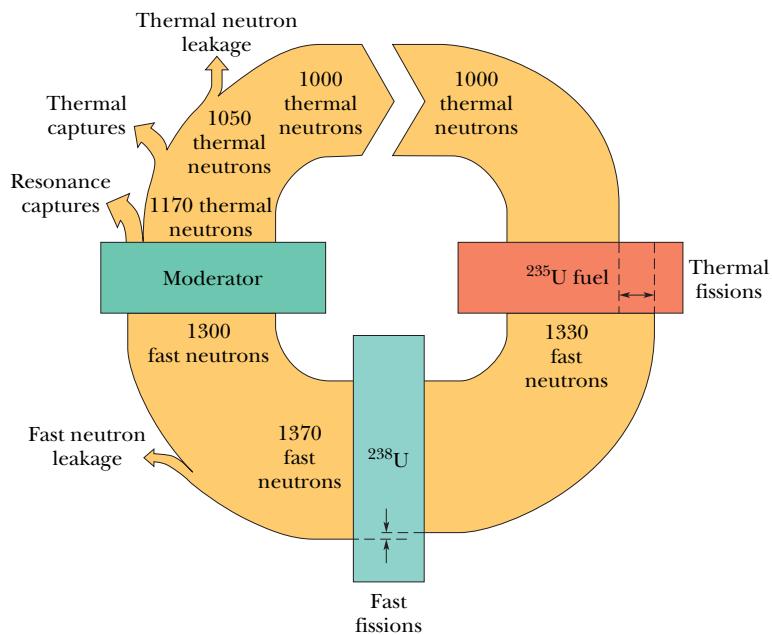


Fig. 43-5 Neutron bookkeeping in a reactor. A generation of 1000 thermal neutrons interacts with the ^{235}U fuel, the ^{238}U matrix, and the moderator. They produce 1370 neutrons by fission, but 370 of these are lost by nonfission capture or by leakage, meaning that 1000 thermal neutrons are left to form the next generation. The figure is drawn for a reactor running at a steady power level.

complete cycle, or *generation*, in the reactor core. They produce 1330 neutrons by fission in the ^{235}U fuel and 40 neutrons by fast fission in ^{238}U , which gives 370 neutrons more than the original 1000, all of them fast. When the reactor is operating at a steady power level, exactly the same number of neutrons (370) is then lost by leakage from the core and by nonfission capture, leaving 1000 thermal neutrons to start the next generation. In this cycle, of course, each of the 370 neutrons produced by fission events represents a deposit of energy in the reactor core, heating up the core.

The *multiplication factor* k —an important reactor parameter—is the ratio of the number of neutrons present at the conclusion of a particular generation to the number present at the beginning of that generation. In Fig. 43-5, the multiplication factor is 1000/1000, or exactly unity. For $k = 1$, the operation of the reactor is said to be exactly *critical*, which is what we wish it to be for steady-power operation. Reactors are actually designed so that they are inherently *supercritical* ($k > 1$); the multiplication factor is then adjusted to critical operation ($k = 1$) by inserting **control rods** into the reactor core. These rods, containing a material such as cadmium that absorbs neutrons readily, can be inserted farther to reduce the operating power level and withdrawn to increase the power level or to compensate for the tendency of reactors to go *subcritical* as (neutron-absorbing) fission products build up in the core during continued operation.

If you pulled out one of the control rods rapidly, how fast would the reactor power level increase? This *response time* is controlled by the fascinating circumstance that a small fraction of the neutrons generated by fission do not escape promptly from the newly formed fission fragments but are emitted from these fragments later, as the fragments decay by beta emission. Of the 370 “new” neutrons produced in Fig. 43-5, for example, perhaps 16 are delayed, being emitted from fragments following beta decays whose half-lives range from 0.2 to 55 s. These delayed neutrons are few in number, but they serve the essential purpose of slowing the reactor response time to match practical mechanical reaction times.

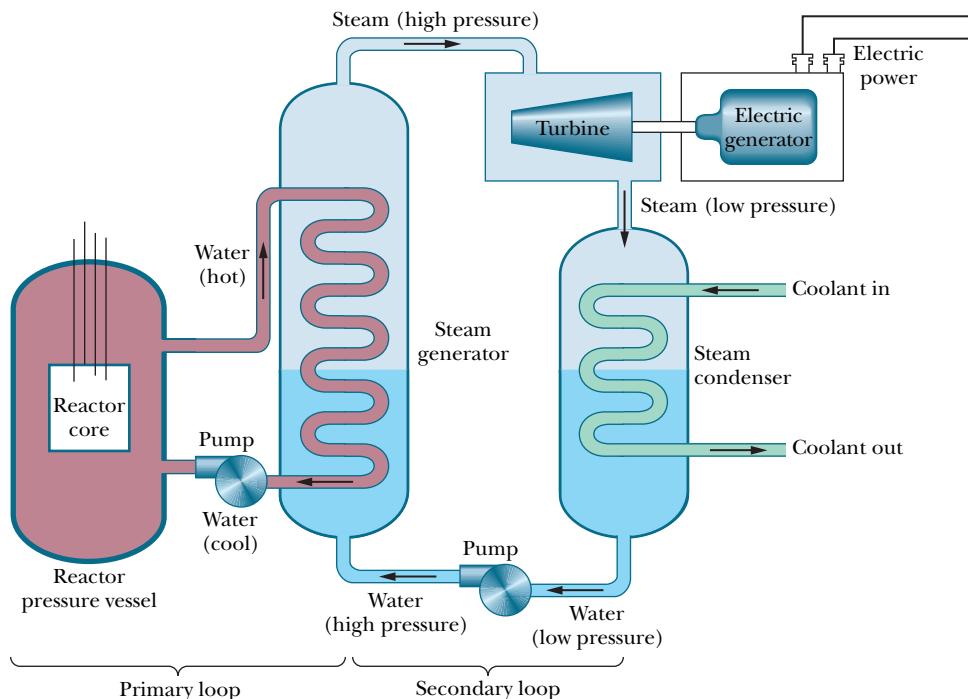
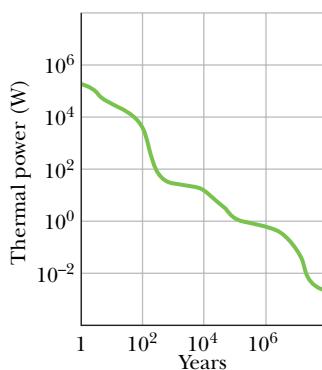


Fig. 43-6 A simplified layout of a nuclear power plant, based on a pressurized-water reactor. Many features are omitted—among them the arrangement for cooling the reactor core in case of an emergency.

Figure 43-6 shows the broad outlines of an electrical power plant based on a *pressurized-water reactor* (PWR), a type in common use in North America. In such a reactor, water is used both as the moderator and as the heat transfer medium. In the *primary loop*, water is circulated through the reactor vessel and transfers energy at high temperature and pressure (possibly 600 K and 150 atm) from the hot reactor core to the steam generator, which is part of the *secondary loop*. In the steam generator, evaporation provides high-pressure steam to operate the turbine that drives the electric generator. To complete the secondary loop, low-pressure steam from the turbine is cooled and condensed to water and forced back into the steam generator by a pump. To give some idea of scale, a typical reactor vessel for a 1000 MW (electric) plant may be 12 m high and weigh 4 MN. Water flows through the primary loop at a rate of about 1 ML/min.

An unavoidable feature of reactor operation is the accumulation of radioactive wastes, including both fission products and heavy *transuranic* nuclides such as plutonium and americium. One measure of their radioactivity is the rate at which they release energy in thermal form. Figure 43-7 shows the thermal power generated by such wastes from one year's operation of a typical large nuclear plant. Note that both scales are logarithmic. Most “spent” fuel rods from power reactor operation are stored on site, immersed in water; permanent secure storage facilities for reactor waste have yet to be completed. Much weapons-derived radioactive waste accumulated during World War II and in subsequent years is also still in on-site storage.

Fig. 43-7 The thermal power released by the radioactive wastes from one year's operation of a typical large nuclear power plant, shown as a function of time. The curve is the superposition of the effects of many radionuclides, with a wide variety of half-lives. Note that both scales are logarithmic.



Sample Problem**Nuclear reactor: efficiency, fission rate, consumption rate**

A large electric generating station is powered by a pressurized-water nuclear reactor. The thermal power produced in the reactor core is 3400 MW, and 1100 MW of electricity is generated by the station. The *fuel charge* is 8.60×10^4 kg of uranium, in the form of uranium oxide, distributed among 5.70×10^4 fuel rods. The uranium is enriched to 3.0% ^{235}U .

(a) What is the station's efficiency?

KEY IDEA

The efficiency for this power plant or any other energy device is given by this: Efficiency is the ratio of the output power (rate at which useful energy is provided) to the input power (rate at which energy must be supplied).

Calculation: Here the efficiency (eff) is

$$\begin{aligned} \text{eff} &= \frac{\text{useful output}}{\text{energy input}} = \frac{1100 \text{ MW (electric)}}{3400 \text{ MW (thermal)}} \\ &= 0.32, \text{ or } 32\%. \end{aligned} \quad (\text{Answer})$$

The efficiency—as for all power plants—is controlled by the second law of thermodynamics. To run this plant, energy at the rate of 3400 MW – 1100 MW, or 2300 MW, must be discharged as thermal energy to the environment.

(b) At what rate R do fission events occur in the reactor core?

KEY IDEAS

1. The fission events provide the input power P of 3400 MW ($= 3.4 \times 10^9$ J/s).
2. From Eq. 43-6, the energy Q released by each event is about 200 MeV.

Calculation: For steady-state operation (P is constant), we find

$$\begin{aligned} R &= \frac{P}{Q} = \left(\frac{3.4 \times 10^9 \text{ J/s}}{200 \text{ MeV/fission}} \right) \left(\frac{1 \text{ MeV}}{1.60 \times 10^{-13} \text{ J}} \right) \\ &= 1.06 \times 10^{20} \text{ fissions/s} \\ &\approx 1.1 \times 10^{20} \text{ fissions/s}. \end{aligned} \quad (\text{Answer})$$

(c) At what rate (in kilograms per day) is the ^{235}U fuel disappearing? Assume conditions at start-up.

KEY IDEA

^{235}U disappears due to two processes: (1) the fission process with the rate calculated in part (b) and (2) the nonfission capture of neutrons at about one-fourth that rate.

Calculations: The total rate at which the number of atoms of ^{235}U decreases is

$$(1 + 0.25)(1.06 \times 10^{20} \text{ atoms/s}) = 1.33 \times 10^{20} \text{ atoms/s}.$$

We want the corresponding decrease in the mass of the ^{235}U fuel. We start with the mass of each ^{235}U atom. We cannot use the molar mass for uranium listed in Appendix F because that molar mass is for ^{238}U , the most common uranium isotope. Instead, we shall assume that the mass of each ^{235}U atom in atomic mass units is equal to the mass number A . Thus, the mass of each ^{235}U atom is 235 u ($= 3.90 \times 10^{-25}$ kg). Then the rate at which the ^{235}U fuel disappears is

$$\begin{aligned} \frac{dM}{dt} &= (1.33 \times 10^{20} \text{ atoms/s})(3.90 \times 10^{-25} \text{ kg/atom}) \\ &= 5.19 \times 10^{-5} \text{ kg/s} \approx 4.5 \text{ kg/d}. \end{aligned} \quad (\text{Answer})$$

(d) At this rate of fuel consumption, how long would the fuel supply of ^{235}U last?

Calculation: At start-up, we know that the total mass of ^{235}U is 3.0% of the 8.60×10^4 kg of uranium oxide. So, the time T required to consume this total mass of ^{235}U at the steady rate of 4.5 kg/d is

$$T = \frac{(0.030)(8.60 \times 10^4 \text{ kg})}{4.5 \text{ kg/d}} \approx 570 \text{ d}. \quad (\text{Answer})$$

In practice, the fuel rods must be replaced (usually in batches) before their ^{235}U content is entirely consumed.

(e) At what rate is mass being converted to other forms of energy by the fission of ^{235}U in the reactor core?

KEY IDEA

The conversion of mass energy to other forms of energy is linked only to the fissioning that produces the input power (3400 MW) and not to the nonfission capture of neutrons (although both these processes affect the rate at which ^{235}U is consumed).

Calculation: From Einstein's relation $E = mc^2$, we can write

$$\begin{aligned} \frac{dm}{dt} &= \frac{dE/dt}{c^2} = \frac{3.4 \times 10^9 \text{ W}}{(3.00 \times 10^8 \text{ m/s})^2} \\ &= 3.8 \times 10^{-8} \text{ kg/s} = 3.3 \text{ g/d}. \end{aligned} \quad (\text{Answer})$$

We see that the mass conversion rate is about the mass of one common coin per day, considerably less than the fuel consumption rate calculated in (c).



Additional examples, video, and practice available at WileyPLUS

43-5 A Natural Nuclear Reactor

On December 2, 1942, when their reactor first became operational (Fig. 43-8), Enrico Fermi and his associates had every right to assume that they had put into operation the first fission reactor that had ever existed on this planet. About 30 years later it was discovered that, if they did in fact think that, they were wrong.

Some two billion years ago, in a uranium deposit recently mined in Gabon, West Africa, a natural fission reactor apparently went into operation and ran for perhaps several hundred thousand years before shutting down. We can test whether this could actually have happened by considering two questions:

- 1. Was There Enough Fuel?** The fuel for a uranium-based fission reactor must be the easily fissionable isotope ^{235}U , which, as noted earlier, constitutes only 0.72% of natural uranium. This isotopic ratio has been measured for terrestrial samples, in Moon rocks, and in meteorites; in all cases the abundance values are the same. The clue to the discovery in West Africa was that the uranium in that deposit was deficient in ^{235}U , some samples having abundances as low as 0.44%. Investigation led to the speculation that this deficit in ^{235}U could be accounted for if, at some earlier time, the ^{235}U was partially consumed by the operation of a natural fission reactor.

The serious problem remains that, with an isotopic abundance of only 0.72%, a reactor can be assembled (as Fermi and his team learned) only after thoughtful design and with scrupulous attention to detail. There seems no chance that a nuclear reactor could go critical “naturally.”

However, things were different in the distant past. Both ^{235}U and ^{238}U are radioactive, with half-lives of 7.04×10^8 y and 44.7×10^8 y, respectively. Thus, the half-life of the readily fissionable ^{235}U is about 6.4 times shorter than that of ^{238}U . Because ^{235}U decays faster, there was more of it, relative to ^{238}U , in the past. Two billion years ago, in fact, this abundance was not 0.72%, as it is now, but 3.8%. This abundance happens to be just about the abundance to which natural uranium is artificially enriched to serve as fuel in modern power reactors.

With this readily fissionable fuel available, the presence of a natural reactor (provided certain other conditions are met) is less surprising. The fuel was there. Two billion years ago, incidentally, the highest order of life-form to have evolved was the blue-green alga.

- 2. What Is the Evidence?** The mere depletion of ^{235}U in an ore deposit does not prove the existence of a natural fission reactor. One looks for more convincing evidence.

If there was a reactor, there must now be fission products. Of the 30 or so elements whose stable isotopes are produced in a reactor, some must still remain. Study of their isotopic abundances could provide the evidence we need.

Of the several elements investigated, the case of neodymium is spectacularly convincing. Figure 43-9a shows the isotopic abundances of the seven stable neodymium isotopes as they are normally found in nature. Figure 43-9b shows these abundances as they appear among the ultimate stable fission products of the fission of ^{235}U . The clear differences are not surprising, considering the totally different origins of the two sets of isotopes. Note particularly that ^{142}Nd , the dominant isotope in the natural element, is absent from the fission products.

The big question is: What do the neodymium isotopes found in the uranium ore body in West Africa look like? If a natural reactor operated there, we would expect to find isotopes from *both* sources (that is, natural isotopes as well as fission-produced isotopes). Figure 43-9c shows the abundances after dual-source and other corrections have been made to the data. Comparison of Figs. 43-9b and c indicates that there was indeed a natural fission reactor at work.



Fig. 43-8 A painting of the first nuclear reactor, assembled during World War II on a squash court at the University of Chicago by a team headed by Enrico Fermi. This reactor was built of lumps of uranium embedded in blocks of graphite.

(Gary Sheehan, Birth of the Atomic Age, 1957. Reproduced courtesy Chicago Historical Society)

43-6 THERMONUCLEAR FUSION: THE BASIC PROCESS

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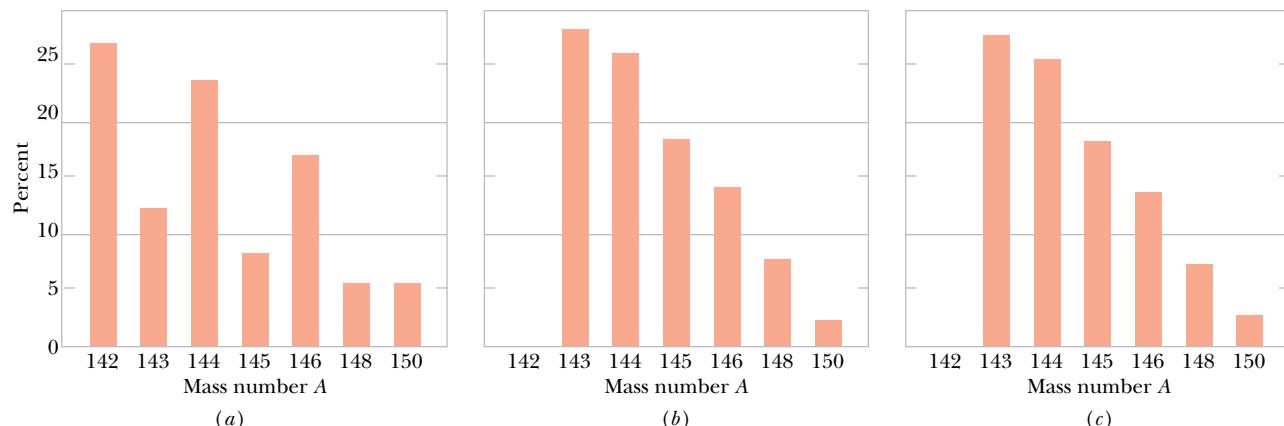


Fig. 43-9 The distribution by mass number of the isotopes of neodymium as they occur in (a) natural terrestrial deposits of the ores of this element and (b) the spent fuel of a power reactor. (c) The distribution (after several corrections) found for neodymium from the uranium mine in Gabon, West Africa. Note that (b) and (c) are virtually identical and are quite different from (a).

43-6 Thermonuclear Fusion: The Basic Process

The binding energy curve of Fig. 42-7 shows that energy can be released if two light nuclei combine to form a single larger nucleus, a process called nuclear **fusion**. That process is hindered by the Coulomb repulsion that acts to prevent the two positively charged particles from getting close enough to be within range of their attractive nuclear forces and thus “fusing.” The height of this *Coulomb barrier* depends on the charges and the radii of the two interacting nuclei. For two protons ($Z = 1$), the barrier height is 400 keV. For more highly charged particles, of course, the barrier is correspondingly higher.

To generate useful amounts of energy, nuclear fusion must occur in bulk matter. The best hope for bringing this about is to raise the temperature of the material until the particles have enough energy—due to their thermal motions alone—to penetrate the Coulomb barrier. We call this process **thermonuclear fusion**.

In thermonuclear studies, temperatures are reported in terms of the kinetic energy K of interacting particles via the relation

$$K = kT, \quad (43-9)$$

in which K is the kinetic energy corresponding to the *most probable speed* of the interacting particles, k is the Boltzmann constant, and the temperature T is in kelvins. Thus, rather than saying, “The temperature at the center of the Sun is 1.5×10^7 K,” it is more common to say, “The temperature at the center of the Sun is 1.3 keV.”

Room temperature corresponds to $K \approx 0.03$ eV; a particle with only this amount of energy could not hope to overcome a barrier as high as, say, 400 keV. Even at the center of the Sun, where $kT = 1.3$ keV, the outlook for thermonuclear fusion does not seem promising at first glance. Yet we know that thermonuclear fusion not only occurs in the core of the Sun but is the dominant feature of that body and of all other stars.

The puzzle is solved when we realize two facts: (1) The energy calculated with Eq. 43-9 is that of the particles with the *most probable* speed, as defined in Section 19-7; there is a long tail of particles with much higher speeds and, correspondingly, much higher energies. (2) The barrier heights that we have calculated represent the *peaks* of the barriers. Barrier tunneling can occur at energies considerably below those peaks, as we saw with alpha decay in Section 42-5.

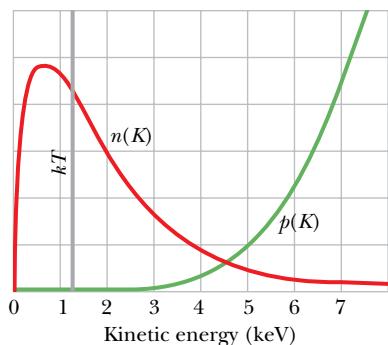


Fig. 43-10 The curve marked $n(K)$ gives the number density per unit energy for protons at the center of the Sun. The curve marked $p(K)$ gives the probability of barrier penetration (and hence fusion) for proton–proton collisions at the Sun’s core temperature. The vertical line marks the value of kT in the core of the Sun is indicated by the vertical line in the figure; note that many of the Sun’s core protons have energies greater than this value.

Figure 43-10 sums things up. The curve marked $n(K)$ in this figure is a Maxwell distribution curve for the protons in the Sun’s core, drawn to correspond to the Sun’s central temperature. This curve differs from the Maxwell distribution curve given in Fig. 19-7 in that here the curve is drawn in terms of energy and not of speed. Specifically, for any kinetic energy K , the expression $n(K) dK$ gives the probability that a proton will have a kinetic energy lying between the values K and $K + dK$. The value of kT in the core of the Sun is indicated by the vertical line in the figure; note that many of the Sun’s core protons have energies greater than this value.

The curve marked $p(K)$ in Fig. 43-10 is the probability of barrier penetration by two colliding protons. The two curves in Fig. 43-10 suggest that there is a particular proton energy at which proton–proton fusion events occur at a maximum rate. At energies much above this value, the barrier is transparent enough but too few protons have these energies, and so the fusion reaction cannot be sustained. At energies much below this value, plenty of protons have these energies but the Coulomb barrier is too formidable.



CHECKPOINT 2

Which of these potential fusion reactions will *not* result in the net release of energy:
 (a) ${}^6\text{Li} + {}^6\text{Li}$, (b) ${}^4\text{He} + {}^4\text{He}$, (c) ${}^{12}\text{C} + {}^{12}\text{C}$, (d) ${}^{20}\text{Ne} + {}^{20}\text{Ne}$, (e) ${}^{35}\text{Cl} + {}^{35}\text{Cl}$, and (f) ${}^{14}\text{N} + {}^{35}\text{Cl}$? (Hint: Consult the curve of Fig. 42-7.)

Sample Problem

Fusion in a gas of protons, and the required temperature

Assume a proton is a sphere of radius $R \approx 1 \text{ fm}$. Two protons are fired at each other with the same kinetic energy K .

- (a) What must K be if the particles are brought to rest by their mutual Coulomb repulsion when they are just “touching” each other? We can take this value of K as a representative measure of the height of the Coulomb barrier.

KEY IDEAS

The mechanical energy E of the two-proton system is conserved as the protons move toward each other and momentarily stop. In particular, the initial mechanical energy E_i is equal to the mechanical energy E_f when they stop. The initial energy E_i consists only of the total kinetic energy $2K$ of the two protons. When the protons stop, energy E_f consists only of the electric potential energy U of the system, as given by Eq. 24-43 ($U = q_1 q_2 / 4\pi\epsilon_0 r$).

Calculations: Here the distance r between the protons when they stop is their center-to-center distance $2R$, and their charges q_1 and q_2 are both e . Then we can write the conservation of energy $E_i = E_f$ as

$$2K = \frac{1}{4\pi\epsilon_0} \frac{e^2}{2R}.$$

This yields, with known values,

$$\begin{aligned} K &= \frac{e^2}{16\pi\epsilon_0 R} \\ &= \frac{(1.60 \times 10^{-19} \text{ C})^2}{(16\pi)(8.85 \times 10^{-12} \text{ F/m})(1 \times 10^{-15} \text{ m})} \\ &= 5.75 \times 10^{-14} \text{ J} = 360 \text{ keV} \approx 400 \text{ keV}. \quad (\text{Answer}) \end{aligned}$$

- (b) At what temperature would a proton in a gas of protons have the average kinetic energy calculated in (a) and thus have energy equal to the height of the Coulomb barrier?

KEY IDEA

If we treat the proton gas as an ideal gas, then from Eq. 19-24, the average energy of the protons is $K_{\text{avg}} = \frac{3}{2}kT$, where k is the Boltzmann constant.

Calculation: Solving that equation for T and using the result of (a) yield

$$\begin{aligned} T &= \frac{2K_{\text{avg}}}{3k} = \frac{(2)(5.75 \times 10^{-14} \text{ J})}{(3)(1.38 \times 10^{-23} \text{ J/K})} \\ &\approx 3 \times 10^9 \text{ K}. \quad (\text{Answer}) \end{aligned}$$

The temperature of the core of the Sun is only about $1.5 \times 10^7 \text{ K}$; thus fusion in the Sun’s core must involve protons whose energies are *far* above the average energy.



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43-7 Thermonuclear Fusion in the Sun and Other Stars

The Sun has been radiating energy at the rate of 3.9×10^{26} W for several billion years. Where does all this energy come from? It does not come from chemical burning. (Even if the Sun were made of coal and had its own oxygen, burning the coal would last only 1000 y.) It also does not come from the Sun shrinking, transferring gravitational potential energy to thermal energy. (Its lifetime would be short by a factor of at least 500.) That leaves only thermonuclear fusion. The Sun, as you will see, burns not coal but hydrogen, and in a nuclear furnace, not an atomic or chemical one.

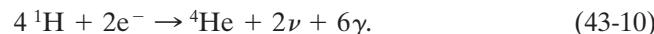
The fusion reaction in the Sun is a multistep process in which hydrogen is burned to form helium, hydrogen being the “fuel” and helium the “ashes.” Figure 43-11 shows the **proton–proton (p-p) cycle** by which this occurs.

The p-p cycle starts with the collision of two protons (${}^1\text{H} + {}^1\text{H}$) to form a deuteron (${}^2\text{H}$), with the simultaneous creation of a positron (e^+) and a neutrino (ν). The positron very quickly encounters a free electron (e^-) in the Sun, and both particles annihilate (see Section 21-6), their mass energy appearing as two gamma-ray photons (γ).

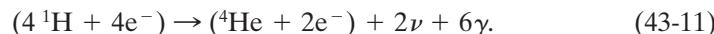
A pair of such events is shown in the top row of Fig. 43-11. These events are actually extremely rare. In fact, only once in about 10^{26} proton–proton collisions is a deuteron formed; in the vast majority of cases, the two protons simply rebound elastically from each other. It is the slowness of this “bottleneck” process that regulates the rate of energy production and keeps the Sun from exploding. In spite of this slowness, there are so very many protons in the huge and dense volume of the Sun’s core that deuterium is produced in just this way at the rate of 10^{12} kg/s.

Once a deuteron has been produced, it quickly collides with another proton and forms a ${}^3\text{He}$ nucleus, as the middle row of Fig. 43-11 shows. Two such ${}^3\text{He}$ nuclei may eventually (within 10^5 y; there is plenty of time) find each other, forming an alpha particle (${}^4\text{He}$) and two protons, as the bottom row in the figure shows.

Overall, we see from Fig. 43-11 that the p-p cycle amounts to the combination of four protons and two electrons to form an alpha particle, two neutrinos, and six gamma-ray photons. That is,



Let us now add two electrons to each side of Eq. 43-10, obtaining



The quantities in the two sets of parentheses then represent *atoms* (not bare nuclei) of hydrogen and of helium. That allows us to compute the energy release in the overall reaction of Eq. 43-10 (and Eq. 43-11) as

$$\begin{aligned} Q &= -\Delta m c^2 \\ &= -[4.002\,603 \text{ u} - (4)(1.007\,825 \text{ u})][931.5 \text{ MeV/u}] \\ &= 26.7 \text{ MeV}, \end{aligned}$$

in which 4.002 603 u is the mass of a helium atom and 1.007 825 u is the mass of a hydrogen atom. Neutrinos have a negligibly small mass, and gamma-ray photons have no mass; thus, they do not enter into the calculation of the disintegration energy.

This same value of Q follows (as it must) from adding up the Q values for the separate steps of the proton–proton cycle in Fig. 43-11. Thus,

$$\begin{aligned} Q &= (2)(0.42 \text{ MeV}) + (2)(1.02 \text{ MeV}) + (2)(5.49 \text{ MeV}) + 12.86 \text{ MeV} \\ &= 26.7 \text{ MeV}. \end{aligned}$$

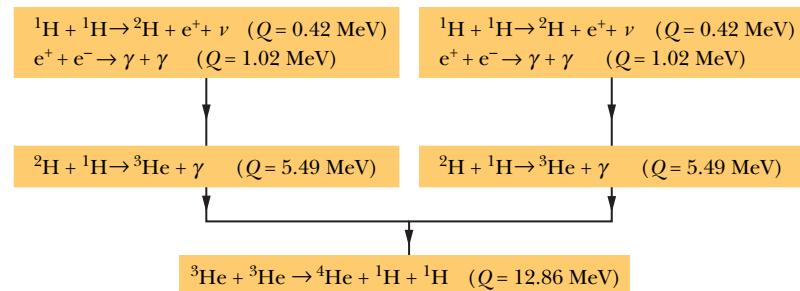


Fig. 43-11 The proton–proton mechanism that accounts for energy production in the Sun. In this process, protons fuse to form an alpha particle (${}^4\text{He}$), with a net energy release of 26.7 MeV for each event.

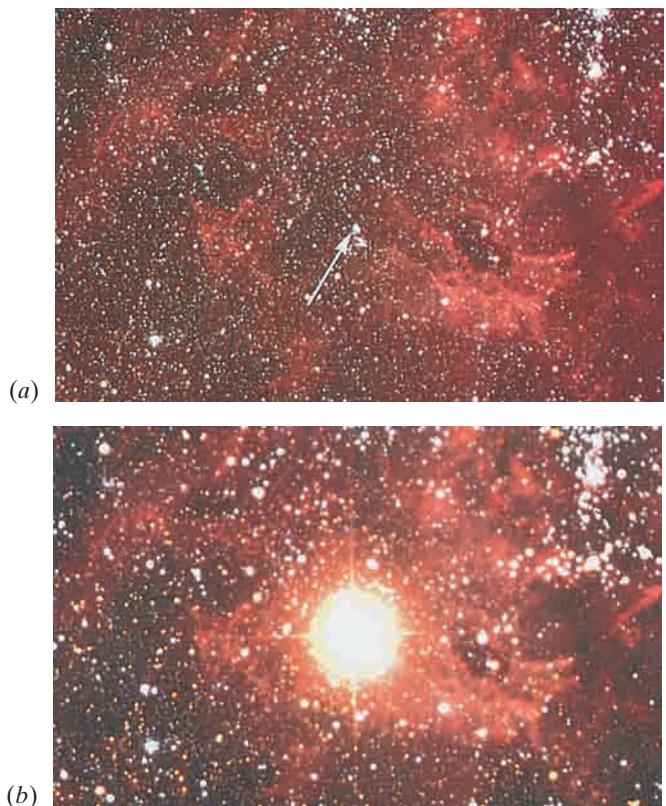


Fig. 43-12 (a) The star known as Sanduleak, as it appeared until 1987. (b) We then began to intercept light from the star's supernova, designated SN1987a; the explosion was 100 million times brighter than our Sun and could be seen with the unaided eye. (Courtesy Anglo Australian Telescope Board)

About 0.5 MeV of this energy is carried out of the Sun by the two neutrinos indicated in Eqs. 43-10 and 43-12; the rest ($= 26.2$ MeV) is deposited in the core of the Sun as thermal energy. That thermal energy is then gradually transported to the Sun's surface, where it is radiated away from the Sun as electromagnetic waves, including visible light.

Hydrogen burning has been going on in the Sun for about 5×10^9 y, and calculations show that there is enough hydrogen left to keep the Sun going for about the same length of time into the future. In 5 billion years, however, the Sun's core, which by that time will be largely helium, will begin to cool and the Sun will start to collapse under its own gravity. This will raise the core temperature and cause the outer envelope to expand, turning the Sun into what is called a *red giant*.

If the core temperature increases to about 10^8 K again, energy can be produced through fusion once more—this time by burning helium to make carbon. As a star evolves further and becomes still hotter, other elements can be formed by other fusion reactions. However, elements more massive than those near the peak of the binding energy curve of Fig. 42-7 cannot be produced by further fusion processes.

Elements with mass numbers beyond the peak are thought to be formed by neutron capture during cataclysmic stellar explosions that we call *supernovas* (Fig. 43-12). In such an event the outer shell of the star is blown outward into space, where it mixes with the tenuous medium that fills the space between the stars. It

is from this medium, continually enriched by debris from stellar explosions, that new stars form, by condensation under the influence of the gravitational force.

The abundance on Earth of elements heavier than hydrogen and helium suggests that our solar system has condensed out of interstellar material that contained the remnants of such explosions. Thus, all the elements around us—including those in our own bodies—were manufactured in the interiors of stars that no longer exist. As one scientist put it: “In truth, we are the children of the stars.”

Sample Problem

Consumption rate of hydrogen in the Sun

At what rate dm/dt is hydrogen being consumed in the core of the Sun by the p-p cycle of Fig. 43-11?

KEY IDEA

The rate dE/dt at which energy is produced by hydrogen (proton) consumption within the Sun is equal to the rate P at which energy is radiated by the Sun:

$$P = \frac{dE}{dt}.$$

Calculations: To bring the mass consumption rate dm/dt into the power equation, we can rewrite it as

$$P = \frac{dE}{dt} = \frac{dE}{dm} \frac{dm}{dt} \approx \frac{\Delta E}{\Delta m} \frac{dm}{dt}, \quad (43-12)$$

where ΔE is the energy produced when protons of mass Δm are consumed. From our discussion in this section, we know that 26.2 MeV ($= 4.20 \times 10^{-12}$ J) of thermal energy is produced when four protons are consumed. That is, $\Delta E = 4.20 \times 10^{-12}$ J for a mass consumption of $\Delta m = 4(1.67 \times 10^{-27}$ kg). Substituting these data into Eq. 43-12 and using the power P of the Sun given in Appendix C, we find that

$$\begin{aligned} \frac{dm}{dt} &= \frac{\Delta m}{\Delta E} P = \frac{4(1.67 \times 10^{-27} \text{ kg})}{4.20 \times 10^{-12} \text{ J}} (3.90 \times 10^{26} \text{ W}) \\ &= 6.2 \times 10^{11} \text{ kg/s.} \end{aligned} \quad (\text{Answer})$$

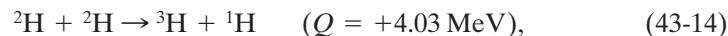
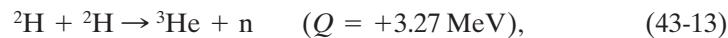
Thus, a huge amount of hydrogen is consumed by the Sun every second. However, you need not worry too much about the Sun running out of hydrogen, because its mass of 2×10^{30} kg will keep it burning for a long, long time.

43-8 Controlled Thermonuclear Fusion

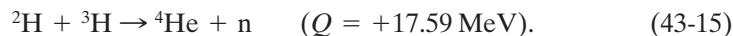
The first thermonuclear reaction on Earth occurred at Eniwetok Atoll on November 1, 1952, when the United States exploded a fusion device, generating an energy release equivalent to 10 million tons of TNT. The high temperatures and densities needed to initiate the reaction were provided by using a fission bomb as a trigger.

A sustained and controllable source of fusion power—a fusion reactor as part of, say, an electric generating plant—is considerably more difficult to achieve. That goal is nonetheless being pursued vigorously in many countries around the world, because many people look to the fusion reactor as the power source of the future, at least for the generation of electricity.

The p-p scheme displayed in Fig. 43-11 is not suitable for an Earth-bound fusion reactor because it is hopelessly slow. The process succeeds in the Sun only because of the enormous density of protons in the center of the Sun. The most attractive reactions for terrestrial use appear to be two deuteron–deuteron (d-d) reactions,



and the deuteron–triton (d-t) reaction



(The nucleus of the hydrogen isotope ${}^3\text{H}$ (tritium) is called the *triton* and has a half-life of 12.3 y.) Deuterium, the source of deuterons for these reactions, has an isotopic abundance of only 1 part in 6700 but is available in unlimited quantities as a component of seawater. Proponents of power from the nucleus have described our ultimate power choice—after we have burned up all our fossil fuels—as either “burning rocks” (fission of uranium extracted from ores) or “burning water” (fusion of deuterium extracted from water).

There are three requirements for a successful thermonuclear reactor:

1. *A High Particle Density n.* The number density of interacting particles (the number of, say, deuterons per unit volume) must be great enough to ensure that the d-d collision rate is high enough. At the high temperatures required, the deuterium would be completely ionized, forming an electrically neutral **plasma** (ionized gas) of deuterons and electrons.
2. *A High Plasma Temperature T.* The plasma must be hot. Otherwise the colliding deuterons will not be energetic enough to penetrate the Coulomb barrier that tends to keep them apart. A plasma ion temperature of 35 keV, corresponding to $4 \times 10^8 \text{ K}$, has been achieved in the laboratory. This is about 30 times higher than the Sun’s central temperature.
3. *A Long Confinement Time τ .* A major problem is containing the hot plasma long enough to maintain it at a density and a temperature sufficiently high to ensure the fusion of enough of the fuel. Because it is clear that no solid container can withstand the high temperatures that are necessary, clever confining techniques are called for; we shall shortly discuss two of them.

It can be shown that, for the successful operation of a thermonuclear reactor using the d-t reaction, it is necessary to have

$$n\tau > 10^{20} \text{ s/m}^3. \quad (43-16)$$

This condition, known as **Lawson’s criterion**, tells us that we have a choice between confining a lot of particles for a short time or fewer particles for a longer time. Also, the plasma temperature must be high enough.

Two approaches to controlled nuclear power generation are currently under study. Although neither approach has yet been successful, both are being pursued

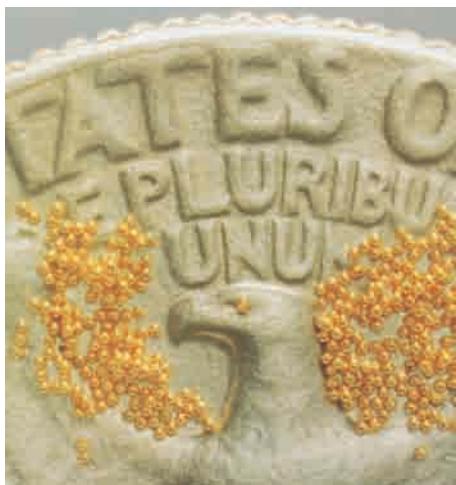


Fig. 43-13 The small spheres on the quarter are deuterium–tritium fuel pellets, designed to be used in a laser fusion chamber. (Courtesy Los Alamos National Laboratory, New Mexico)

because of their promise and because of the potential importance of controlled fusion to solving the world's energy problems.

Magnetic Confinement

One avenue to controlled fusion is to contain the fusing material in a very strong magnetic field—hence the name **magnetic confinement**. In one version of this approach, a suitably shaped magnetic field is used to confine the hot plasma in an evacuated doughnut-shaped chamber called a **tokamak** (the name is an abbreviation consisting of parts of three Russian words). The magnetic forces acting on the charged particles that make up the hot plasma keep the plasma from touching the walls of the chamber.

The plasma is heated by inducing a current in it and by bombarding it with an externally accelerated beam of particles. The first goal of this approach is to achieve **break-even**, which occurs when the Lawson criterion is met or exceeded. The ultimate goal is **ignition**, which corresponds to a self-sustaining thermonuclear reaction and a net generation of energy.

Inertial Confinement

A second approach, called **inertial confinement**, involves “zapping” a solid fuel pellet from all sides with intense laser beams, evaporating some material from the surface of the pellet. This boiled-off material causes an inward-moving shock wave that compresses the core of the pellet, increasing both its particle density and its temperature. The process is called inertial confinement because (a) the fuel is *confined* to the pellet and (b) the particles do not escape from the heated pellet during the very short zapping interval because of their *inertia* (their mass).

Laser fusion, using the inertial confinement approach, is being investigated in many laboratories in the United States and elsewhere. At the Lawrence Livermore Laboratory, for example, deuterium–tritium fuel pellets, each smaller than a grain of sand (Fig. 43-13), are to be zapped by 10 synchronized high-power laser pulses symmetrically arranged around the pellet. The laser pulses are designed to deliver, in total, some 200 kJ of energy to each fuel pellet in less than a nanosecond. This is a delivered power of about 2×10^{14} W during the pulse, which is roughly 100 times the total installed (sustained) electrical power generating capacity of the world!

Sample Problem

Laser fusion: number of particles and Lawson's criterion

Suppose a fuel pellet in a laser fusion device contains equal numbers of deuterium and tritium atoms (and no other material). The density $d = 200 \text{ kg/m}^3$ of the pellet is increased by a factor of 10^3 by the action of the laser pulses.

- (a) How many particles per unit volume (both deuterons and tritons) does the pellet contain in its compressed state? The molar mass M_d of deuterium atoms is $2.0 \times 10^{-3} \text{ kg/mol}$, and the molar mass M_t of tritium atoms is $3.0 \times 10^{-3} \text{ kg/mol}$.

KEY IDEA

For a system consisting of only one type of particle, we can write the (mass) density (the mass per unit volume) of the

system in terms of the particle masses and number density (the number of particles per unit volume):

$$\left(\begin{array}{c} \text{density,} \\ \text{kg/m}^3 \end{array} \right) = \left(\begin{array}{c} \text{number density,} \\ \text{m}^{-3} \end{array} \right) \left(\begin{array}{c} \text{particle mass,} \\ \text{kg} \end{array} \right). \quad (43-17)$$

Let n be the total number of particles per unit volume in the compressed pellet. Then, because we know that the device contains equal numbers of deuterium and tritium atoms, the number of deuterium atoms per unit volume is $n/2$, and the number of tritium atoms per unit volume is also $n/2$.

Calculations: We can extend Eq. 43-17 to the system consisting of the two types of particles by writing the density d^* of

the compressed pellet as the sum of the individual densities:

$$d^* = \frac{n}{2} m_d + \frac{n}{2} m_t, \quad (43-18)$$

where m_d and m_t are the masses of a deuterium atom and a tritium atom, respectively. We can replace those masses with the given molar masses by substituting

$$m_d = \frac{M_d}{N_A} \quad \text{and} \quad m_t = \frac{M_t}{N_A},$$

where N_A is Avogadro's number. After making those replacements and substituting $1000d$ for the compressed density d^* , we solve Eq. 43-18 for n to obtain

$$n = \frac{2000dN_A}{M_d + M_t},$$

which gives us



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$$\begin{aligned} n &= \frac{(2000)(200 \text{ kg/m}^3)(6.02 \times 10^{23} \text{ mol}^{-1})}{2.0 \times 10^{-3} \text{ kg/mol} + 3.0 \times 10^{-3} \text{ kg/mol}} \\ &= 4.8 \times 10^{31} \text{ m}^{-3}. \end{aligned} \quad (\text{Answer})$$

- (b) According to Lawson's criterion, how long must the pellet maintain this particle density if breakeven operation is to take place?

KEY IDEA

If breakeven operation is to occur, the compressed density must be maintained for a time period τ given by Eq. 43-16 ($n\tau > 10^{20} \text{ s/m}^3$).

Calculation: We can now write

$$\tau > \frac{10^{20} \text{ s/m}^3}{4.8 \times 10^{31} \text{ m}^{-3}} \approx 10^{-12} \text{ s.} \quad (\text{Answer})$$

(The plasma temperature must also be suitably high.)

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REVIEW & SUMMARY

Energy from the Nucleus Nuclear processes are about a million times more effective, per unit mass, than chemical processes in transforming mass into other forms of energy.

Nuclear Fission Equation 43-1 shows a **fission** of ^{236}U induced by thermal neutrons bombarding ^{235}U . Equations 43-2 and 43-3 show the beta-decay chains of the primary fragments. The energy released in such a fission event is $Q \approx 200 \text{ MeV}$.

Fission can be understood in terms of the collective model, in which a nucleus is likened to a charged liquid drop carrying a certain excitation energy. A potential barrier must be tunneled through if fission is to occur. The ability of a nucleus to undergo fission depends on the relationship between the barrier height E_b and the excitation energy E_n .

The neutrons released during fission make possible a **fission chain reaction**. Figure 43-5 shows the neutron balance for one cycle of a typical reactor. Figure 43-6 suggests the layout of a complete nuclear power plant.

Nuclear Fusion The release of energy by the **fusion** of two

light nuclei is inhibited by their mutual Coulomb barrier (due to the electric repulsion between the two collections of protons). Fusion can occur in bulk matter only if the temperature is high enough (that is, if the particle energy is high enough) for appreciable barrier tunneling to occur.

The Sun's energy arises mainly from the thermonuclear burning of hydrogen to form helium by the **proton–proton cycle** outlined in Fig. 43-11. Elements up to $A \approx 56$ (the peak of the binding energy curve) can be built up by other fusion processes once the hydrogen fuel supply of a star has been exhausted.

Controlled Fusion Controlled **thermonuclear fusion** for energy generation has not yet been achieved. The d-d and d-t reactions are the most promising mechanisms. A successful fusion reactor must satisfy **Lawson's criterion**,

$$n\tau > 10^{20} \text{ s/m}^3, \quad (43-16)$$

and must have a suitably high plasma temperature T .

In a **tokamak** the plasma is confined by a magnetic field. In **laser fusion** inertial confinement is used.

QUESTIONS

- 1 In the fission process



what number goes in (a) the elevated box (the superscript) and (b) the descended box (the value of Z)?

- 2 If a fusion process requires an absorption of energy, does the average binding energy per nucleon increase or decrease?

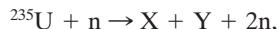
- 3 Suppose a ^{238}U nucleus "swallows" a neutron and then decays not by fission but by beta-minus decay, in which it emits an electron and a neutrino. Which nuclide remains after this decay: ^{239}Pu , ^{238}Np , ^{239}Np , or ^{238}Pa ?

- 4 Do the initial fragments formed by fission have more protons than neutrons, more neutrons than protons, or about the same number of each?

** View All Solutions Here **

1214 CHAPTER 43 ENERGY FROM THE NUCLEUS

5 For the fission reaction



rank the following possibilities for X (or Y), most likely first: ^{152}Nd , ^{140}I , ^{128}In , ^{115}Pd , ^{105}Mo . (Hint: See Fig. 43-1.)

6 To make the newly discovered, very large elements of the periodic table, researchers shoot a medium-size nucleus at a large nucleus. Sometimes a projectile nucleus and a target nucleus fuse to form one of the very large elements. In such a fusion, is the mass of the product greater than or less than the sum of the masses of the projectile and target nuclei?

7 If we split a nucleus into two smaller nuclei, with a release of energy, has the average binding energy per nucleon increased or decreased?

8 Which of these elements is *not* “cooked up” by thermonuclear fusion processes in stellar interiors: carbon, silicon, chromium, bromine?

9 Lawson's criterion for the d-t reaction (Eq. 43-16) is $n\tau > 10^{20} \text{ s/m}^3$. For the d-d reaction, do you expect the number on the right-hand side to be the same, smaller, or larger?

10 About 2% of the energy generated in the Sun's core by the p-p reaction is carried out of the Sun by neutrinos. Is the energy associated with this neutrino flux equal to, greater than, or less than the energy radiated from the Sun's surface as electromagnetic radiation?

11 A nuclear reactor is operating at a certain power level, with its multiplication factor k adjusted to unity. If the control rods are used to reduce the power output of the reactor to 25% of its former value, is the multiplication factor now a little less than unity, substantially less than unity, or still equal to unity?

12 Pick the most likely member of each pair to be one of the initial fragments formed by a fission event: (a) ^{93}Sr or ^{93}Ru , (b) ^{140}Gd or ^{140}I , (c) ^{155}Nd or ^{155}Lu . (Hint: See Fig. 42-5 and the periodic table, and consider the neutron abundance.)

PROBLEMS



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty



Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>

sec. 43-3 A Model for Nuclear Fission

•1 The isotope ^{235}U decays by alpha emission with a half-life of $7.0 \times 10^8 \text{ y}$. It also decays (rarely) by spontaneous fission, and if the alpha decay did not occur, its half-life due to spontaneous fission alone would be $3.0 \times 10^{17} \text{ y}$. (a) At what rate do spontaneous fission decays occur in 1.0 g of ^{235}U ? (b) How many ^{235}U alpha-decay events are there for every spontaneous fission event?

•2 The nuclide ^{238}Np requires 4.2 MeV for fission. To remove a neutron from this nuclide requires an energy expenditure of 5.0 MeV. Is ^{237}Np fissionable by thermal neutrons?

•3 A thermal neutron (with approximately zero kinetic energy) is absorbed by a ^{238}U nucleus. How much energy is transferred from mass energy to the resulting oscillation of the nucleus? Here are some atomic masses and the neutron mass.

^{237}U	237.048 723 u	^{238}U	238.050 782 u
^{239}U	239.054 287 u	^{240}U	240.056 585 u
n	1.008 664 u		

•4 The fission properties of the plutonium isotope ^{239}Pu are very similar to those of ^{235}U . The average energy released per fission is 180 MeV. How much energy, in MeV, is released if all the atoms in 1.00 kg of pure ^{239}Pu undergo fission?

•5 During the Cold War, the Premier of the Soviet Union threatened the United States with 2.0 megaton ^{239}Pu warheads. (Each would have yielded the equivalent of an explosion of 2.0 megatons of TNT, where 1 megaton of TNT releases $2.6 \times 10^{28} \text{ MeV}$.) If the plutonium that actually fissioned had been 8.00% of the total mass of the plutonium in such a warhead, what was that total mass?

•6 (a)–(d) Complete the following table, which refers to the generalized fission reaction $^{235}\text{U} + \text{n} \rightarrow \text{X} + \text{Y} + bn$.

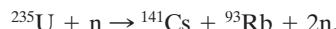
X	Y	b
^{140}Xe	(a)	1
^{139}I	(b)	2
(c)	^{100}Zr	2
^{141}Cs	^{92}Rb	(d)

•7 At what rate must ^{235}U nuclei undergo fission by neutron bombardment to generate energy at the rate of 1.0 W? Assume that $Q = 200 \text{ MeV}$.

•8 (a) Calculate the disintegration energy Q for the fission of ^{98}Mo into two equal parts. The masses you will need are 97.905 41 u for ^{98}Mo and 48.950 02 u for ^{49}Sc . (b) If Q turns out to be positive, discuss why this process does not occur spontaneously.

•9 (a) How many atoms are contained in 1.0 kg of pure ^{235}U ? (b) How much energy, in joules, is released by the complete fissioning of 1.0 kg of ^{235}U ? Assume $Q = 200 \text{ MeV}$. (c) For how long would this energy light a 100 W lamp?

•10 Calculate the energy released in the fission reaction



Here are some atomic and particle masses.

^{235}U	235.043 92 u	^{93}Rb	92.921 57 u
^{141}Cs	140.919 63 u	n	1.008 66 u

•11 Calculate the disintegration energy Q for the fission of ^{52}Cr into two equal fragments. The masses you will need are



- 12** Consider the fission of ^{238}U by fast neutrons. In one fission event, no neutrons are emitted and the final stable end products, after the beta decay of the primary fission fragments, are ^{140}Ce and ^{99}Ru . (a) What is the total of the beta-decay events in the two beta-decay chains? (b) Calculate Q for this fission process. The relevant atomic and particle masses are

^{238}U	238.050 79 u	^{140}Ce	139.905 43 u
n	1.008 66 u	^{99}Ru	98.905 94 u

- 13** Assume that immediately after the fission of ^{236}U according to Eq. 43-1, the resulting ^{140}Xe and ^{94}Sr nuclei are just touching at their surfaces. (a) Assuming the nuclei to be spherical, calculate the electric potential energy associated with the repulsion between the two fragments. (*Hint:* Use Eq. 42-3 to calculate the radii of the fragments.) (b) Compare this energy with the energy released in a typical fission event.

- 14** A ^{236}U nucleus undergoes fission and breaks into two middle-mass fragments, ^{140}Xe and ^{96}Sr . (a) By what percentage does the surface area of the fission products differ from that of the original ^{236}U nucleus? (b) By what percentage does the volume change? (c) By what percentage does the electric potential energy change? The electric potential energy of a uniformly charged sphere of radius r and charge Q is given by

$$U = \frac{3}{5} \left(\frac{Q^2}{4\pi\epsilon_0 r} \right).$$

- 15** **SSM** A 66 kiloton atomic bomb is fueled with pure ^{235}U (Fig. 43-14), 4.0% of which actually undergoes fission. (a) What is the mass of the uranium in the bomb? (It is not 66 kilotons—that is the amount of released energy specified in terms of the mass of TNT required to produce the same amount of energy.) (b) How many primary fission fragments are produced? (c) How many fission neutrons generated are released to the environment? (On average, each fission produces 2.5 neutrons.)



Fig. 43-14 Problem 15. A “button” of ^{235}U ready to be recast and machined for a warhead. (Courtesy Martin Marietta Energy Systems/U.S. Department of Energy)

- 16** In an atomic bomb, energy release is due to the uncontrolled fission of plutonium ^{239}Pu (or ^{235}U). The bomb’s rating is the magnitude of the released energy, specified in terms of the mass of TNT

required to produce the same energy release. One megaton of TNT releases 2.6×10^{28} MeV of energy. (a) Calculate the rating, in tons of TNT, of an atomic bomb containing 95.0 kg of ^{239}Pu , of which 2.5 kg actually undergoes fission. (See Problem 4.) (b) Why is the other 92.5 kg of ^{239}Pu needed if it does not fission?

- 17** **SSM** **WWW** In a particular fission event in which ^{235}U is fissioned by slow neutrons, no neutron is emitted and one of the primary fission fragments is ^{83}Ge . (a) What is the other fragment? The disintegration energy is $Q = 170$ MeV. How much of this energy goes to (b) the ^{83}Ge fragment and (c) the other fragment? Just after the fission, what is the speed of (d) the ^{83}Ge fragment and (e) the other fragment?

sec. 43-4 The Nuclear Reactor

- 18** A 200 MW fission reactor consumes half its fuel in 3.00 y. How much ^{235}U did it contain initially? Assume that all the energy generated arises from the fission of ^{235}U and that this nuclide is consumed only by the fission process.

- 19** The neutron generation time t_{gen} in a reactor is the average time needed for a fast neutron emitted in one fission event to be slowed to thermal energies by the moderator and then initiate another fission event. Suppose the power output of a reactor at time $t = 0$ is P_0 . Show that the power output a time t later is $P(t)$, where $P(t) = P_0 k^{t/t_{\text{gen}}}$ and k is the multiplication factor. For constant power output, $k = 1$.

- 20** A reactor operates at 400 MW with a neutron generation time (see Problem 19) of 30.0 ms. If its power increases for 5.00 min with a multiplication factor of 1.0003, what is the power output at the end of the 5.00 min?

- 21** The thermal energy generated when radiation from radionuclides is absorbed in matter can serve as the basis for a small power source for use in satellites, remote weather stations, and other isolated locations. Such radionuclides are manufactured in abundance in nuclear reactors and may be separated chemically from the spent fuel. One suitable radionuclide is ^{238}Pu ($T_{1/2} = 87.7$ y), which is an alpha emitter with $Q = 5.50$ MeV. At what rate is thermal energy generated in 1.00 kg of this material?

- 22** The neutron generation time t_{gen} (see Problem 19) in a particular reactor is 1.0 ms. If the reactor is operating at a power level of 500 MW, about how many free neutrons are present in the reactor at any moment?

- 23** **SSM** **WWW** The neutron generation time (see Problem 19) of a particular reactor is 1.3 ms. The reactor is generating energy at the rate of 1200.0 MW. To perform certain maintenance checks, the power level must temporarily be reduced to 350.00 MW. It is desired that the transition to the reduced power level take 2.6000 s. To what (constant) value should the multiplication factor be set to effect the transition in the desired time?

- 24** (See Problem 21.) Among the many fission products that may be extracted chemically from the spent fuel of a nuclear reactor is ^{90}Sr ($T_{1/2} = 29$ y). This isotope is produced in typical large reactors at the rate of about 18 kg/y. By its radioactivity, the isotope generates thermal energy at the rate of 0.93 W/g. (a) Calculate the effective disintegration energy Q_{eff} associated with the decay of a ^{90}Sr nucleus. (This energy Q_{eff} includes contributions from the decay of the ^{90}Sr daughter products in its decay chain but not from neutrinos, which escape totally from the sample.) (b) It is desired to construct a power source generating 150 W (electric power) to use in operating electronic equipment in an underwater acoustic

1216 CHAPTER 43 ENERGY FROM THE NUCLEUS

beacon. If the power source is based on the thermal energy generated by ${}^{90}\text{Sr}$ and if the efficiency of the thermal–electric conversion process is 5.0%, how much ${}^{90}\text{Sr}$ is needed?

- 25 **SSM** (a) A neutron of mass m_n and kinetic energy K makes a head-on elastic collision with a stationary atom of mass m . Show that the fractional kinetic energy loss of the neutron is given by

$$\frac{\Delta K}{K} = \frac{4m_n m}{(m + m_n)^2}.$$

Find $\Delta K/K$ for each of the following acting as the stationary atom: (b) hydrogen, (c) deuterium, (d) carbon, and (e) lead. (f) If $K = 1.00 \text{ MeV}$ initially, how many such head-on collisions would it take to reduce the neutron's kinetic energy to a thermal value (0.025 eV) if the stationary atoms it collides with are deuterium, a commonly used moderator? (In actual moderators, most collisions are not head-on.)

sec. 43-5 A Natural Nuclear Reactor

- 26 How long ago was the ratio ${}^{235}\text{U}/{}^{238}\text{U}$ in natural uranium deposits equal to 0.15?

- 27 The natural fission reactor discussed in Section 43-5 is estimated to have generated 15 gigawatt-years of energy during its lifetime. (a) If the reactor lasted for 200 000 y, at what average power level did it operate? (b) How many kilograms of ${}^{235}\text{U}$ did it consume during its lifetime?

- 28 Some uranium samples from the natural reactor site described in Section 43-5 were found to be slightly *enriched* in ${}^{235}\text{U}$, rather than depleted. Account for this in terms of neutron absorption by the abundant isotope ${}^{238}\text{U}$ and the subsequent beta and alpha decay of its products.

- 29 **SSM** The uranium ore mined today contains only 0.72% of fissionable ${}^{235}\text{U}$, too little to make reactor fuel for thermal-neutron fission. For this reason, the mined ore must be enriched with ${}^{235}\text{U}$. Both ${}^{235}\text{U}$ ($T_{1/2} = 7.0 \times 10^8 \text{ y}$) and ${}^{238}\text{U}$ ($T_{1/2} = 4.5 \times 10^9 \text{ y}$) are radioactive. How far back in time would natural uranium ore have been a practical reactor fuel, with a ${}^{235}\text{U}/{}^{238}\text{U}$ ratio of 3.0%?

sec. 43-6 Thermonuclear Fusion: The Basic Process

- 30 Verify that the fusion of 1.0 kg of deuterium by the reaction



could keep a 100 W lamp burning for $2.5 \times 10^4 \text{ y}$.

- 31 **SSM** Calculate the height of the Coulomb barrier for the head-on collision of two deuterons, with effective radius 2.1 fm.

- 32 For overcoming the Coulomb barrier for fusion, methods other than heating the fusible material have been suggested. For example, if you were to use two particle accelerators to accelerate two beams of deuterons directly toward each other so as to collide head-on, (a) what voltage would each accelerator require in order for the colliding deuterons to overcome the Coulomb barrier? (b) Why do you suppose this method is not presently used?

- 33 Calculate the Coulomb barrier height for two ${}^7\text{Li}$ nuclei that are fired at each other with the same initial kinetic energy K . (*Hint:* Use Eq. 42-3 to calculate the radii of the nuclei.)

- 34 In Fig. 43-10, the equation for $n(K)$, the number density per unit energy for particles, is

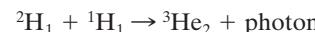
$$n(K) = 1.13n \frac{K^{1/2}}{(kT)^{3/2}} e^{-K/kT},$$

where n is the total particle number density. At the center of the Sun, the temperature is $1.50 \times 10^7 \text{ K}$ and the average proton energy K_{avg} is 1.94 keV. Find the ratio of the proton number density at 5.00 keV to the number density at the average proton energy.

sec. 43-7 Thermonuclear Fusion in the Sun and Other Stars

- 35 Assume that the protons in a hot ball of protons each have a kinetic energy equal to kT , where k is the Boltzmann constant and T is the absolute temperature. If $T = 1 \times 10^7 \text{ K}$, what (approximately) is the least separation any two protons can have?

- 36 **GO** What is the Q of the following fusion process?



Here are some atomic masses.

${}^2\text{H}_1$	2.014 102 u	${}^1\text{H}_1$	1.007 825 u
${}^3\text{He}_2$	3.016 029 u		

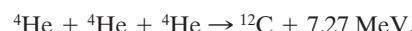
- 37 The Sun has mass $2.0 \times 10^{30} \text{ kg}$ and radiates energy at the rate $3.9 \times 10^{26} \text{ W}$. (a) At what rate is its mass changing? (b) What fraction of its original mass has it lost in this way since it began to burn hydrogen, about $4.5 \times 10^9 \text{ y}$ ago?

- 38 We have seen that Q for the overall proton–proton fusion cycle is 26.7 MeV. How can you relate this number to the Q values for the reactions that make up this cycle, as displayed in Fig. 43-11?

- 39 **GO** Show that the energy released when three alpha particles fuse to form ${}^{12}\text{C}$ is 7.27 MeV. The atomic mass of ${}^4\text{He}$ is 4.0026 u, and that of ${}^{12}\text{C}$ is 12.0000 u.

- 40 Calculate and compare the energy released by (a) the fusion of 1.0 kg of hydrogen deep within the Sun and (b) the fission of 1.0 kg of ${}^{235}\text{U}$ in a fission reactor.

- 41 **GO** A star converts all its hydrogen to helium, achieving a 100% helium composition. Next it converts the helium to carbon via the triple-alpha process,



The mass of the star is $4.6 \times 10^{32} \text{ kg}$, and it generates energy at the rate of $5.3 \times 10^{30} \text{ W}$. How long will it take to convert all the helium to carbon at this rate?

- 42 Verify the three Q values reported for the reactions given in Fig. 43-11. The needed atomic and particle masses are

${}^1\text{H}$	1.007 825 u	${}^4\text{He}$	4.002 603 u
${}^2\text{H}$	2.014 102 u	e^\pm	0.000 548 6 u
${}^3\text{He}$	3.016 029 u		

(*Hint:* Distinguish carefully between atomic and nuclear masses, and take the positrons properly into account.)

- 43 Figure 43-15 shows an early proposal for a hydrogen bomb. The fusion fuel is deuterium, ${}^2\text{H}$. The high temperature and particle density needed for fusion are provided by an atomic bomb “trigger” that involves a ${}^{235}\text{U}$ or ${}^{239}\text{Pu}$ fission fuel arranged to impress an imploding, compressive shock wave on the deuterium. The fusion reaction is

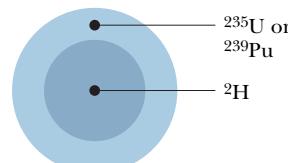
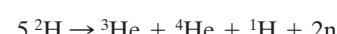


Fig. 43-15 Problem 43.



(a) Calculate Q for the fusion reaction. For needed atomic masses, see Problem 42. (b) Calculate the rating (see Problem 16) of the fusion part of the bomb if it contains 500 kg of deuterium, 30.0% of which undergoes fusion.

••44 Assume that the core of the Sun has one-eighth of the Sun's mass and is compressed within a sphere whose radius is one-fourth of the solar radius. Assume further that the composition of the core is 35% hydrogen by mass and that essentially all the Sun's energy is generated there. If the Sun continues to burn hydrogen at the current rate of 6.2×10^{11} kg/s, how long will it be before the hydrogen is entirely consumed? The Sun's mass is 2.0×10^{30} kg.

••45 (a) Calculate the rate at which the Sun generates neutrinos. Assume that energy production is entirely by the proton–proton fusion cycle. (b) At what rate do solar neutrinos reach Earth?

••46 In certain stars the *carbon cycle* is more effective than the proton–proton cycle in generating energy. This carbon cycle is

$$\begin{aligned} {}^{12}\text{C} + {}^1\text{H} &\rightarrow {}^{13}\text{N} + \gamma, & Q_1 &= 1.95 \text{ MeV}, \\ {}^{13}\text{N} &\rightarrow {}^{13}\text{C} + e^+ + \nu, & Q_2 &= 1.19, \\ {}^{13}\text{C} + {}^1\text{H} &\rightarrow {}^{14}\text{N} + \gamma, & Q_3 &= 7.55, \\ {}^{14}\text{N} + {}^1\text{H} &\rightarrow {}^{15}\text{O} + \gamma, & Q_4 &= 7.30, \\ {}^{15}\text{O} &\rightarrow {}^{15}\text{N} + e^+ + \nu, & Q_5 &= 1.73, \\ {}^{15}\text{N} + {}^1\text{H} &\rightarrow {}^{12}\text{C} + {}^4\text{He}, & Q_6 &= 4.97. \end{aligned}$$

(a) Show that this cycle is exactly equivalent in its overall effects to the proton–proton cycle of Fig. 43-11. (b) Verify that the two cycles, as expected, have the same Q value.

••47 SSM WWW Coal burns according to the reaction $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$. The heat of combustion is 3.3×10^7 J/kg of atomic carbon consumed. (a) Express this in terms of energy per carbon atom. (b) Express it in terms of energy per kilogram of the initial reactants, carbon and oxygen. (c) Suppose that the Sun (mass = 2.0×10^{30} kg) were made of carbon and oxygen in combustible proportions and that it continued to radiate energy at its present rate of 3.9×10^{26} W. How long would the Sun last?

sec. 43-8 Controlled Thermonuclear Fusion

•48 Verify the Q values reported in Eqs. 43-13, 43-14, and 43-15. The needed masses are

${}^1\text{H}$	1.007 825 u	${}^4\text{He}$	4.002 603 u
${}^2\text{H}$	2.014 102 u	n	1.008 665 u
${}^3\text{H}$	3.016 049 u		

••49 Roughly 0.0150% of the mass of ordinary water is due to "heavy water," in which one of the two hydrogens in an H_2O molecule is replaced with deuterium, ${}^2\text{H}$. How much average fusion power could be obtained if we "burned" all the ${}^2\text{H}$ in 1.00 liter of

water in 1.00 day by somehow causing the deuterium to fuse via the reaction ${}^2\text{H} + {}^2\text{H} \rightarrow {}^3\text{He} + n$?

Additional Problems

50 The effective Q for the proton–proton cycle of Fig. 43-11 is 26.2 MeV. (a) Express this as energy per kilogram of hydrogen consumed. (b) The power of the Sun is 3.9×10^{26} W. If its energy derives from the proton–proton cycle, at what rate is it losing hydrogen? (c) At what rate is it losing mass? (d) Account for the difference in the results for (b) and (c). (e) The mass of the Sun is 2.0×10^{30} kg. If it loses mass at the constant rate calculated in (c), how long will it take to lose 0.10% of its mass?

51 Many fear that nuclear power reactor technology will increase the likelihood of nuclear war because reactors can be used not only to produce electrical energy but also, as a by-product through neutron capture with inexpensive ${}^{238}\text{U}$, to make ${}^{239}\text{Pu}$, which is a "fuel" for nuclear bombs. What simple series of reactions involving neutron capture and beta decay would yield this plutonium isotope?

52 In the deuteron–triton fusion reaction of Eq. 43-15, what is the kinetic energy of (a) the alpha particle and (b) the neutron? Neglect the relatively small kinetic energies of the two combining particles.

53 Verify that, as stated in Section 43-2, neutrons in equilibrium with matter at room temperature, 300 K, have an average kinetic energy of about 0.04 eV.

54 Verify that, as reported in Table 43-1, fissioning of the ${}^{235}\text{U}$ in 1.0 kg of UO_2 (enriched so that ${}^{235}\text{U}$ is 3.0% of the total uranium) could keep a 100 W lamp burning for 690 y.

55 At the center of the Sun, the density is 1.5×10^5 kg/m³ and the composition is essentially 35% hydrogen by mass and 65% helium by mass. (a) What is the density of protons there? (b) What is the ratio of this density to the density of particles in an ideal gas at standard temperature (0°C) and pressure (1.01×10^5 Pa)?

56 Expressions for the Maxwell speed distribution for molecules in a gas are given in Chapter 19. (a) Show that the *most probable energy* is given by

$$K_p = \frac{1}{2}kT.$$

Verify this result with the energy distribution curve of Fig. 43-10, for which $T = 1.5 \times 10^7$ K. (b) Show that the *most probable speed* is given by

$$v_p = \sqrt{\frac{2kT}{m}}.$$

Find its value for protons at $T = 1.5 \times 10^7$ K. (c) Show that the *energy corresponding to the most probable speed* (which is not the same as the most probable energy) is

$$K_{v,p} = kT.$$

Locate this quantity on the curve of Fig. 43-10.

44

QUARKS, LEPTONS, AND THE BIG BANG

44-1 WHAT IS PHYSICS?

Physicists often refer to the theories of relativity and quantum physics as “modern physics,” to distinguish them from the theories of Newtonian mechanics and Maxwellian electromagnetism, which are lumped together as “classical physics.” As the years go by, the word “modern” seems less and less appropriate for theories whose foundations were laid down in the opening years of the 20th century. After all, Einstein published his paper on the photoelectric effect and his first paper on special relativity in 1905, Bohr published his quantum model of the hydrogen atom in 1913, and Schrödinger published his matter wave equation in 1926. Nevertheless, the label of “modern physics” hangs on.

In this closing chapter we consider two lines of investigation that are truly “modern” but at the same time have the most ancient of roots. They center around two deceptively simple questions:

What is the universe made of?

How did the universe come to be the way it is?

Progress in answering these questions has been rapid in the last few decades.

Many new insights are based on experiments carried out with large particle accelerators. However, as they bang particles together at higher and higher energies using larger and larger accelerators, physicists come to realize that no conceivable Earth-bound accelerator can generate particles with energies great enough to test the ultimate theories of physics. There has been only one source of particles with these energies, and that was the universe itself within the first millisecond of its existence.

In this chapter you will encounter a host of new terms and a veritable flood of particles with names that you should not try to remember. If you are temporarily bewildered, you are sharing the bewilderment of the physicists who lived through these developments and who at times saw nothing but increasing complexity with little hope of understanding. If you stick with it, however, you will come to share the excitement physicists felt as marvelous new accelerators poured out new results, as the theorists put forth ideas each more daring than the last, and as clarity finally sprang from obscurity. The main message of this book is that, although we know a lot about the physics of the world, grand mysteries remain.

44-2 Particles, Particles, Particles

In the 1930s, there were many scientists who thought that the problem of the ultimate structure of matter was well on the way to being solved. The atom could be understood in terms of only three particles—the electron, the proton, and the neutron. Quantum physics accounted well for the structure of the atom and for radioactive alpha decay. The neutrino had been postulated and, although not yet observed, had been incorporated by Enrico Fermi into a successful theory of beta decay. There was hope that quantum theory applied to protons and neutrons would soon account for the structure of the nucleus. What else was there?

The euphoria did not last. The end of that same decade saw the beginning of a period of discovery of new particles that continues to this day. The new particles have names and symbols such as *muon* (μ), *pion* (π), *kaon* (K), and *sigma* (Σ). All the new particles are unstable; that is, they spontaneously transform into other types of particles according to the same functions of time that apply to unstable nuclei. Thus, if N_0 particles of any one type are present in a sample at time $t = 0$, the number N of those particles present at some later time t is given by Eq. 42-15,

$$N = N_0 e^{-\lambda t}, \quad (44-1)$$

the rate of decay R , from an initial value of R_0 , is given by Eq. 42-16,

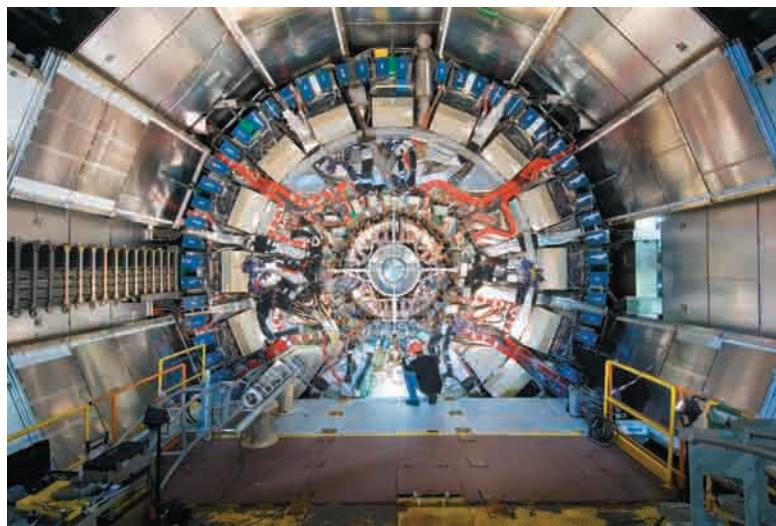
$$R = R_0 e^{-\lambda t}, \quad (44-2)$$

and the half-life $T_{1/2}$, decay constant λ , and mean life τ are related by Eq. 42-18,

$$T_{1/2} = \frac{\ln 2}{\lambda} = \tau \ln 2. \quad (44-3)$$

The half-lives of the new particles range from about 10^{-6} s to 10^{-23} s. Indeed, some of the particles last so briefly that they cannot be detected directly but can only be inferred from indirect evidence.

These new particles are commonly produced in head-on collisions between protons or electrons accelerated to high energies in accelerators at places like Brookhaven National Laboratory (on Long Island, New York), Fermilab (near Chicago), CERN (near Geneva, Switzerland), SLAC (at Stanford University in California), and DESY (near Hamburg, Germany). They are discovered with particle detectors that have grown in sophistication until they rival the size and complexity of entire accelerators of only a few decades ago.



One of the detectors at the Large Hadron Collider at CERN, where the Standard Model of the elementary particles will be put to the test. (© CERN Geneva)

Today there are several hundred known particles. Naming them has strained the resources of the Greek alphabet, and most are known only by an assigned number in a periodically issued compilation. To make sense of this array of particles, we look for simple physical criteria by which we can place the particles in categories. The result is known as the **Standard Model** of particles. Although this model is continuously challenged by theorists, it remains our best scheme of understanding all the particles discovered to date.

To explore the Standard Model, we make the following three rough cuts among the known particles: fermion or boson, hadron or lepton, particle or antiparticle? Let's now look at the categories one by one.

Fermion or Boson?

All particles have an intrinsic angular momentum called **spin**, as we discussed for electrons, protons, and neutrons in Section 32-7. Generalizing the notation of that section, we can write the component of spin \vec{S} in any direction (assume the component to be along a z axis) as

$$S_z = m_s \hbar \quad \text{for } m_s = s, s - 1, \dots, -s, \quad (44-4)$$

in which \hbar is $h/2\pi$, m_s is the *spin magnetic quantum number*, and s is the *spin quantum number*. This last can have either positive half-integer values ($\frac{1}{2}, \frac{3}{2}, \dots$) or nonnegative integer values (0, 1, 2, \dots). For example, an electron has the value $s = \frac{1}{2}$. Hence the spin of an electron (measured along any direction, such as the z direction) can have the values

$$S_z = \frac{1}{2}\hbar \quad (\text{spin up})$$

or $S_z = -\frac{1}{2}\hbar \quad (\text{spin down}).$

Confusingly, the term *spin* is used in two ways: It properly means a particle's intrinsic angular momentum \vec{S} , but it is often used loosely to mean the particle's spin quantum number s . In the latter case, for example, an electron is said to be a spin- $\frac{1}{2}$ particle.

Particles with half-integer spin quantum numbers (like electrons) are called **fermions**, after Fermi, who (simultaneously with Paul Dirac) discovered the statistical laws that govern their behavior. Like electrons, protons and neutrons also have $s = \frac{1}{2}$ and are fermions.

Particles with zero or integer spin quantum numbers are called **bosons**, after Indian physicist Satyendra Nath Bose, who (simultaneously with Albert Einstein) discovered the governing statistical laws for *those* particles. Photons, which have $s = 1$, are bosons; you will soon meet other particles in this class.

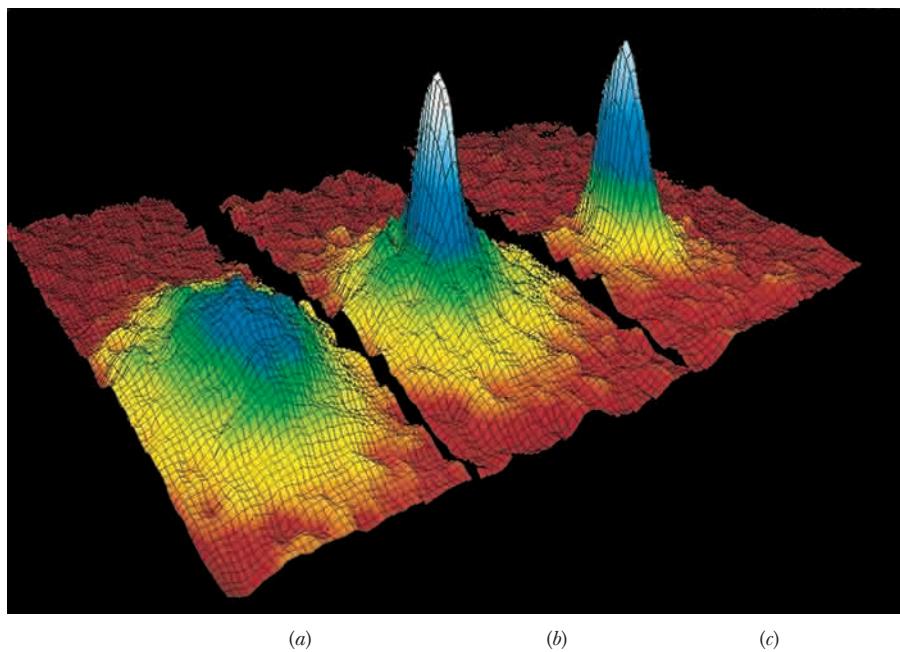
This may seem a trivial way to classify particles, but it is very important for this reason:



Fermions obey the Pauli exclusion principle, which asserts that only a single particle can be assigned to a given quantum state. Bosons *do not* obey this principle. Any number of bosons can occupy a given quantum state.

We saw how important the Pauli exclusion principle is when we “built up” the atoms by assigning (spin- $\frac{1}{2}$) electrons to individual quantum states. Using that principle led to a full accounting of the structure and properties of atoms of different types and of solids such as metals and semiconductors.

Because bosons *do not* obey the Pauli principle, those particles tend to pile up in the quantum state of lowest energy. In 1995 a group in Boulder, Colorado, succeeded in producing a condensate of about 2000 rubidium-87 atoms—they are bosons—in a single quantum state of approximately zero energy.



(a)

(b)

(c)

Fig. 44-1 Three plots of the particle speed distribution in a vapor of rubidium-87 atoms. The temperature of the vapor is successively reduced from plot (a) to plot (c). Plot (c) shows a sharp peak centered around zero speed; that is, all the atoms are in the same quantum state. The achievement of such a Bose–Einstein condensate, often called the Holy Grail of atomic physics, was finally recorded in 1995.
(Courtesy Michael Mathews)

For this to happen, the rubidium has to be a vapor with a temperature so low and a density so great that the de Broglie wavelengths of the individual atoms are greater than the average separation between the atoms. When this condition is met, the wave functions of the individual atoms overlap and the entire assembly becomes a single quantum system (one big atom) called a *Bose–Einstein condensate*. Figure 44-1 shows that, as the temperature of the rubidium vapor is lowered to about 1.70×10^{-7} K, the atoms do indeed “collapse” into a single sharply defined state corresponding to approximately zero speed.

Hadron or Lepton?

We can also classify particles in terms of the four fundamental forces that act on them. The *gravitational force* acts on *all* particles, but its effects at the level of subatomic particles are so weak that we need not consider that force (at least not in today’s research). The *electromagnetic force* acts on all *electrically charged* particles; its effects are well known, and we can take them into account when we need to; we largely ignore this force in this chapter.

We are left with the *strong force*, which is the force that binds nucleons together, and the *weak force*, which is involved in beta decay and similar processes. The weak force acts on all particles, the strong force only on some.

We can, then, roughly classify particles on the basis of whether the strong force acts on them. Particles on which the strong force acts are called **hadrons**. Particles on which the strong force does *not* act, leaving the weak force as the dominant force, are called **leptons**. Protons, neutrons, and pions are hadrons; electrons and neutrinos are leptons.

We can make a further distinction among the hadrons because some of them are bosons (we call them **mesons**); the pion is an example. The other hadrons are fermions (we call them **baryons**); the proton is an example.

Particle or Antiparticle?

In 1928 Dirac predicted that the electron e^- should have a positively charged counterpart of the same mass and spin. The counterpart, the *positron* e^+ , was discovered in cosmic radiation in 1932 by Carl Anderson. Physicists then gradually realized that *every* particle has a corresponding **antiparticle**. The members of such pairs have the same mass and spin but opposite signs of electric charge (if they are charged) and opposite signs of quantum numbers that we have not yet discussed.

At first, *particle* was used to refer to the common particles such as electrons, protons, and neutrons, and *antiparticle* referred to their rarely detected counterparts. Later, for the less common particles, the assignment of *particle* and *antiparticle* was made so as to be consistent with certain conservation laws that we shall discuss later in this chapter. (Confusingly, both particles and antiparticles are sometimes called particles when no distinction is needed.) We often, but not always, represent an antiparticle by putting a bar over the symbol for the particle. Thus, p is the symbol for the proton, and \bar{p} (pronounced “ p bar”) is the symbol for the antiproton.

When a particle meets its antiparticle, the two can *annihilate* each other. That is, the particle and antiparticle disappear and their combined energies reappear in other forms. For an electron annihilating with a positron, this energy reappears as two gamma-ray photons:

$$e^- + e^+ \rightarrow \gamma + \gamma. \quad (44-5)$$

If the electron and positron are stationary when they annihilate, their total energy is their total mass energy, and that energy is then shared equally by the two photons. To conserve momentum and because photons cannot be stationary, the photons fly off in opposite directions.

Large numbers of antihydrogen atoms (each with an antiproton and positron instead of a proton and electron in a hydrogen atom) are now being manufactured and studied at CERN. The Standard Model predicts that a transition in an antihydrogen atom (say, between the first excited state and the ground state) is identical to the same transition in a hydrogen atom. Thus, any difference in the transitions would clearly signal that the Standard Model is erroneous; no difference has yet been spotted.

An assembly of antiparticles, such as an antihydrogen atom, is often called *antimatter* to distinguish it from an assembly of common particles (*matter*). (The terms can easily be confusing when the word “matter” is used to describe anything that has mass.) We can speculate that future scientists and engineers may construct objects of antimatter. However, no evidence suggests that nature has already done this on an astronomical scale because all stars and galaxies appear to consist largely of matter and not antimatter. This is a perplexing observation because it means that when the universe began, some feature biased the conditions toward matter and away from antimatter. (For example, electrons are common but positrons are not.) This bias is still not well understood.

44-3 An Interlude

Before pressing on with the task of classifying the particles, let us step aside for a moment and capture some of the spirit of particle research by analyzing a typical particle event—namely, that shown in the bubble-chamber photograph of Fig. 44-2a.

The tracks in this figure consist of bubbles formed along the paths of electrically charged particles as they move through a chamber filled with liquid

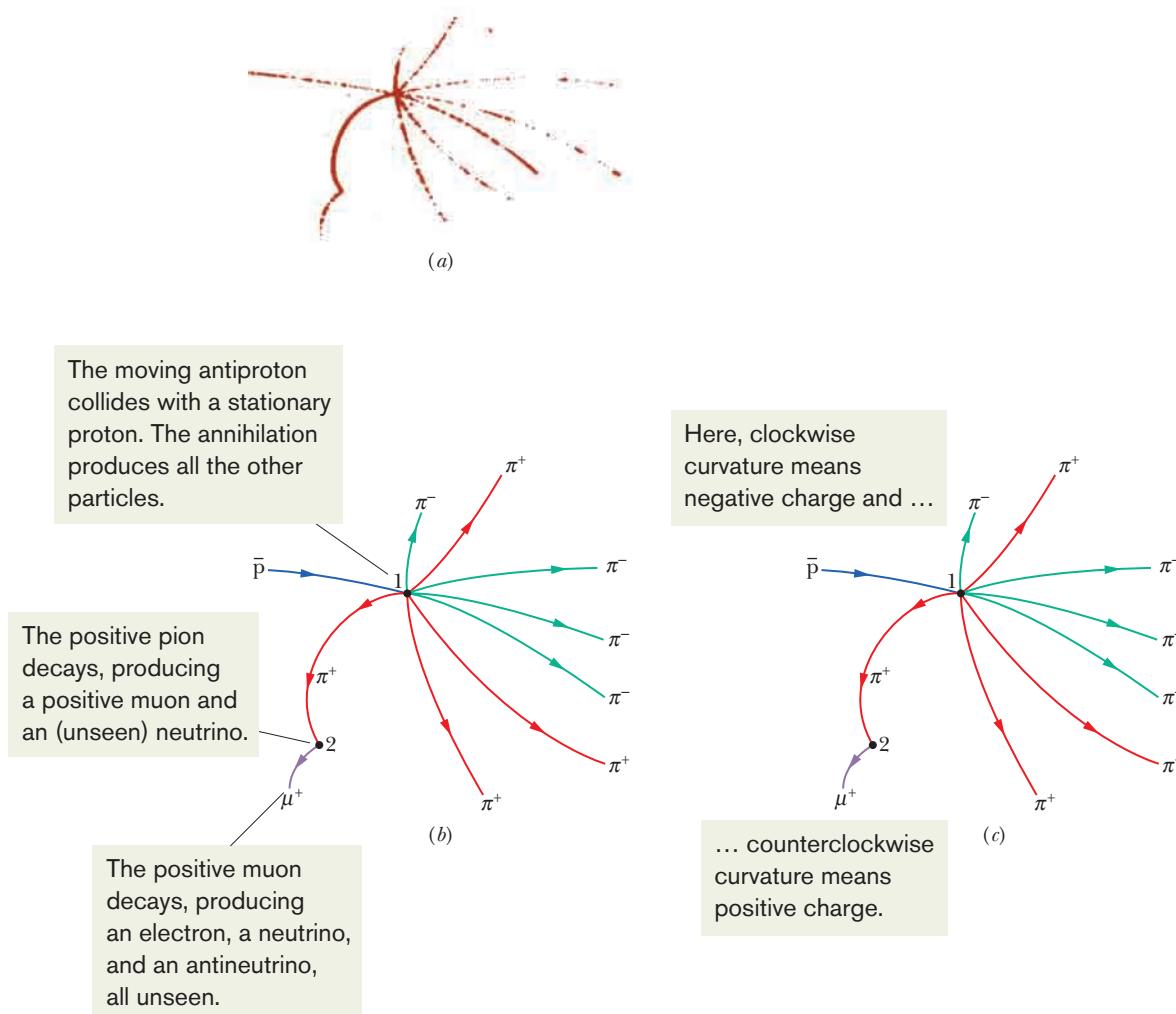


Fig. 44-2 (a) A bubble-chamber photograph of a series of events initiated by an antiproton that enters the chamber from the left. (b) The tracks redrawn and labeled for clarity. (c) The tracks are curved because a magnetic field present in the chamber exerts a deflecting force on each moving charged particle. (Courtesy Lawrence Berkeley Laboratory)

hydrogen. We can identify the particle that makes a particular track by—among other means—measuring the relative spacing between the bubbles. The chamber lies in a uniform magnetic field that deflects the tracks of positively charged particles counterclockwise and the tracks of negatively charged particles clockwise. By measuring the radius of curvature of a track, we can calculate the momentum of the particle that made it. Table 44-1 shows some properties of the particles and antiparticles that participated in the event of Fig. 44-2a, including those that did not make tracks. Following common practice, we express the masses of the particles listed in Table 44-1—and in all other tables in this chapter—in the unit MeV/c^2 . The reason for this notation is that the rest energy of a particle is needed more often than its mass. Thus, the mass of a proton

Table 44-1

The Particles or Antiparticles Involved in the Event of Fig. 44-2

Particle	Symbol	Charge q	Mass (MeV/ c^2)	Spin Quantum Number s	Identity	Mean Life (s)	Antiparticle
Neutrino	ν	0	$\approx 1 \times 10^{-7}$	$\frac{1}{2}$	Lepton	Stable	$\bar{\nu}$
Electron	e^-	-1	0.511	$\frac{1}{2}$	Lepton	Stable	e^+
Muon	μ^-	-1	105.7	$\frac{1}{2}$	Lepton	2.2×10^{-6}	μ^+
Pion	π^+	+1	139.6	0	Meson	2.6×10^{-8}	π^-
Proton	p	+1	938.3	$\frac{1}{2}$	Baryon	Stable	\bar{p}

is shown in Table 44-1 to be 938.3 MeV/ c^2 . To find the proton's rest energy, multiply this mass by c^2 to obtain 938.3 MeV.

The general tools used for the analysis of photographs like Fig. 44-2a are the laws of conservation of energy, linear momentum, angular momentum, and electric charge, along with other conservation laws that we have not yet discussed. Figure 44-2a is actually one of a stereo pair of photographs so that, in practice, these analyses are carried out in three dimensions.

The event of Fig. 44-2a is triggered by an energetic antiproton (\bar{p}) that, generated in an accelerator at the Lawrence Berkeley Laboratory, enters the chamber from the left. There are three separate subevents; one occurs at point 1 in Fig. 44-2b, the second occurs at point 2, and the third occurs out of the frame of the figure. Let's examine each:

1. *Proton–Antiproton Annihilation.* At point 1 in Fig. 44-2b, the initiating antiproton (blue track) slams into a proton of the liquid hydrogen in the chamber, and the result is mutual annihilation. We can tell that annihilation occurred while the incoming antiproton was in flight because most of the particles generated in the encounter move in the forward direction—that is, toward the right in Fig. 44-2. From the principle of conservation of linear momentum, the incoming antiproton must have had a forward momentum when it underwent annihilation. Further, because the particles are charged and moving through a magnetic field, the curvature of the paths reveal whether the particles are negatively charged (like the incident antiproton) or positively charged (Fig. 44-2c).

The total energy involved in the collision of the antiproton and the proton is the sum of the antiproton's kinetic energy and the two (identical) rest energies of those two particles (2×938.3 MeV, or 1876.6 MeV). This is enough energy to create a number of lighter particles and give them kinetic energy. In this case, the annihilation produces four positive pions (red tracks in Fig. 44-2b) and four negative pions (green tracks). (For simplicity, we assume that no gamma-ray photons, which would leave no tracks because they lack electric charge, are produced.) Thus we conclude that the annihilation process is

$$p + \bar{p} \rightarrow 4\pi^+ + 4\pi^- \quad (44-6)$$

We see from Table 44-1 that the positive pions (π^+) are *particles* and the negative pions (π^-) are *antiparticles*. The reaction of Eq. 44-6 is a *strong interaction* (it involves the strong force) because all the particles involved are hadrons.

Let us check whether electric charge is conserved in the reaction. To do so, we can write the electric charge of a particle as qe , in which q is a **charge quantum number**. Then determining whether electric charge is conserved in a process amounts to determining whether the initial net charge quantum

number is equal to the final net charge quantum number. In the process of Eq. 44-6, the initial net charge number is $1 + (-1)$, or 0, and the final net charge number is $4(1) + 4(-1)$, or 0. Thus, charge is conserved.

For the energy balance, note from above that the energy available from the $p-p\bar{p}$ annihilation process is at least the sum of the proton and antiproton rest energies, 1876.6 MeV. The rest energy of a pion is 139.6 MeV, which means the rest energies of the eight pions amount to 8×139.6 MeV, or 1116.8 MeV. This leaves at least about 760 MeV to distribute among the eight pions as kinetic energy. Thus, the requirement of energy conservation is easily met.

- 2. Pion Decay.** Pions are unstable particles and decay with a mean lifetime of 2.6×10^{-8} s. At point 2 in Fig. 44-2b, one of the positive pions comes to rest in the chamber and decays spontaneously into an antimuon μ^+ (purple track) and a neutrino ν :



The neutrino, being uncharged, leaves no track. Both the antimuon and the neutrino are leptons; that is, they are particles on which the strong force does not act. Thus, the decay process of Eq. 44-7, which is governed by the weak force, is described as a *weak interaction*.

Let's consider the energies in the decay. From Table 44-1, the rest energy of an antimuon is 105.7 MeV and the rest energy of a neutrino is approximately 0. Because the pion is at rest when it decays, its energy is just its rest energy, 139.6 MeV. Thus, an energy of 139.6 MeV – 105.7 MeV, or 33.9 MeV, is available to share between the antimuon and the neutrino as kinetic energy.

Let us check whether spin angular momentum is conserved in the process of Eq. 44-7. This amounts to determining whether the net component S_z of spin angular momentum along some arbitrary z axis can be conserved by the process. The spin quantum numbers s of the particles in the process are 0 for the pion π^+ and $\frac{1}{2}$ for both the antimuon μ^+ and the neutrino ν . Thus, for π^+ , the component S_z must be $0\hbar$, and for μ^+ and ν , it can be either $+\frac{1}{2}\hbar$ or $-\frac{1}{2}\hbar$.

The net component S_z is conserved by the process of Eq. 44-7 if there is any way in which the initial S_z ($= 0\hbar$) can be equal to the final net S_z . We see that if one of the products, either μ^+ or ν , has $S_z = +\frac{1}{2}\hbar$ and the other has $S_z = -\frac{1}{2}\hbar$, then their final net value is $0\hbar$. Thus, because S_z can be conserved, the decay process of Eq. 44-7 can occur.

From Eq. 44-7, we also see that the net charge is conserved by the process: before the process the net charge quantum number is +1, and after the process it is $+1 + 0 = +1$.

- 3. Muon Decay.** Muons (whether μ^- or μ^+) are also unstable, decaying with a mean life of 2.2×10^{-6} s. Although the decay products are not shown in Fig. 44-2, the antimuon produced in the reaction of Eq. 44-7 comes to rest and decays spontaneously according to



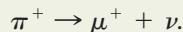
The rest energy of the antimuon is 105.7 MeV, and that of the positron is only 0.511 MeV, leaving 105.2 MeV to be shared as kinetic energy among the three particles produced in the decay process of Eq. 44-8.

You may wonder: Why *two* neutrinos in Eq. 44-8? Why not just one, as in the pion decay in Eq. 44-7? One answer is that the spin quantum numbers of the antimuon, the positron, and the neutrino are each $\frac{1}{2}$; with only one neutrino, the net component S_z of spin angular momentum could not be conserved in the antimuon decay of Eq. 44-8. In Section 44-4 we shall discuss another reason.

Sample Problem

Momentum and kinetic energy in a pion decay

A stationary positive pion can decay according to



What is the kinetic energy of the antimuon μ^+ ? What is the kinetic energy of the neutrino?

KEY IDEA

The pion decay process must conserve both total energy and total linear momentum.

Energy conservation: Let us first write the conservation of total energy (rest energy mc^2 plus kinetic energy K) for the decay process as

$$m_\pi c^2 + K_\pi = m_\mu c^2 + K_\mu + m_\nu c^2 + K_\nu.$$

Because the pion was stationary, its kinetic energy K_π is zero. Then, using the masses listed for m_π , m_μ , and m_ν in Table 44-1, we find

$$\begin{aligned} K_\mu + K_\nu &= m_\pi c^2 - m_\mu c^2 - m_\nu c^2 \\ &= 139.6 \text{ MeV} - 105.7 \text{ MeV} - 0 \\ &= 33.9 \text{ MeV}, \end{aligned} \quad (44-9)$$

where we have approximated m_ν as zero.

Momentum conservation: We cannot solve Eq. 44-9 for either K_μ or K_ν separately, and so let us next apply the principle of conservation of linear momentum to the decay process. Because the pion is stationary when it decays, that principle requires that the muon and neutrino move in opposite directions after the decay. Assume that their motion is along an axis. Then, for components along that axis, we can write the conservation of linear momentum for the decay as

$$p_\pi = p_\mu + p_\nu,$$

which, with $p_\pi = 0$, gives us

$$p_\mu = -p_\nu. \quad (44-10)$$

Relating p and K : We want to relate these momenta p_μ and $-p_\nu$ to the kinetic energies K_μ and K_ν so that we can solve for the kinetic energies. Because we have no reason to believe that classical physics can be applied, we use Eq. 37-54, the momentum–kinetic energy relation from special relativity:

$$(pc)^2 = K^2 + 2Kmc^2. \quad (44-11)$$

From Eq. 44-10, we know that

$$(p_\mu c)^2 = (p_\nu c)^2. \quad (44-12)$$

Substituting from Eq. 44-11 for each side of Eq. 44-12 yields

$$K_\mu^2 + 2K_\mu m_\mu c^2 = K_\nu^2 + 2K_\nu m_\nu c^2.$$

Approximating the neutrino mass to be $m_\nu = 0$, substituting $K_\nu = 33.9 \text{ MeV} - K_\mu$ from Eq. 44-9, and then solving for K_μ , we find

$$\begin{aligned} K_\mu &= \frac{(33.9 \text{ MeV})^2}{(2)(33.9 \text{ MeV} + m_\mu c^2)} \\ &= \frac{(33.9 \text{ MeV})^2}{(2)(33.9 \text{ MeV} + 105.7 \text{ MeV})} \\ &= 4.12 \text{ MeV}. \end{aligned} \quad (\text{Answer})$$

The kinetic energy of the neutrino is then, from Eq. 44-9,

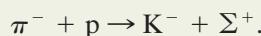
$$\begin{aligned} K_\nu &= 33.9 \text{ MeV} - K_\mu = 33.9 \text{ MeV} - 4.12 \text{ MeV} \\ &= 29.8 \text{ MeV}. \end{aligned} \quad (\text{Answer})$$

We see that, although the magnitudes of the momenta of the two recoiling particles are the same, the neutrino gets the larger share (88%) of the kinetic energy.

Sample Problem

Q in a proton-pion reaction

The protons in the material filling a bubble chamber are bombarded with energetic antiparticles known as negative pions. At collision points, a proton and a pion transform into a negative kaon and a positive sigma:



The rest energies of these particles are

$$\pi^- \quad 139.6 \text{ MeV} \quad K^- \quad 493.7 \text{ MeV}$$

$$p \quad 938.3 \text{ MeV} \quad \Sigma^+ \quad 1189.4 \text{ MeV}$$

What is the Q of the reaction?

KEY IDEA

The Q of a reaction is

$$Q = \left(\begin{array}{c} \text{initial total} \\ \text{mass energy} \end{array} \right) - \left(\begin{array}{c} \text{final total} \\ \text{mass energy} \end{array} \right).$$

Calculation: For the given reaction, we find

$$\begin{aligned} Q &= (m_\pi c^2 + m_p c^2) - (m_K c^2 + m_\Sigma c^2) \\ &= (139.6 \text{ MeV} + 938.3 \text{ MeV}) \\ &\quad - (493.7 \text{ MeV} + 1189.4 \text{ MeV}) \\ &= -605 \text{ MeV}. \end{aligned} \quad (\text{Answer})$$

The minus sign means that the reaction is *endothermic*; that is, the incoming pion (π^-) must have a kinetic energy greater than a certain threshold value if the reaction is to occur. The threshold energy is actually greater than 605 MeV because linear momentum must be conserved. (The incoming pion

has momentum.) This means that the kaon (K^-) and the sigma (Σ^+) not only must be created but also must be given some kinetic energy. A relativistic calculation whose details are beyond our scope shows that the threshold energy for the reaction is 907 MeV.



Additional examples, video, and practice available at WileyPLUS

44-4 The Leptons

In this and the next section, we discuss some of the particles of one of our classification schemes: lepton or hadron. We begin with the leptons, those particles on which the strong force does *not* act. So far, we have encountered the familiar electron and the neutrino that accompanies it in beta decay. The muon, whose decay is described in Eq. 44-8, is another member of this family. Physicists gradually learned that the neutrino that appears in Eq. 44-7, associated with the production of a muon, is *not the same particle* as the neutrino produced in beta decay, associated with the appearance of an electron. We call the former the **muon neutrino** (symbol ν_μ) and the latter the **electron neutrino** (symbol ν_e) when it is necessary to distinguish between them.

These two types of neutrino are known to be different particles because, if a beam of muon neutrinos (produced from pion decay as in Eq. 44-7) strikes a solid target, *only muons*—and never electrons—are produced. On the other hand, if electron neutrinos (produced by the beta decay of fission products in a nuclear reactor) strike a solid target, *only electrons*—and never muons—are produced.

Another lepton, the **tau**, was discovered at SLAC in 1975; its discoverer, Martin Perl, shared the 1995 Nobel Prize in physics. The tau has its own associated neutrino, different still from the other two. Table 44-2 lists all the leptons (both particles and antiparticles); all have a spin quantum number s of $\frac{1}{2}$.

There are reasons for dividing the leptons into three families, each consisting of a particle (electron, muon, or tau), its associated neutrino, and the corresponding antiparticles. Furthermore, there are reasons to believe that there are *only* the three families of leptons shown in Table 44-2. Leptons have no internal structure and no measurable dimensions; they are believed to be truly pointlike fundamental particles when they interact with other particles or with electromagnetic waves.

Table 44-2

The Leptons^a

Family	Particle	Symbol	Mass (MeV/c ²)	Charge q	Antiparticle
Electron	Electron	e^-	0.511	-1	e^+
	Electron neutrino ^b	ν_e	$\approx 1 \times 10^{-7}$	0	$\bar{\nu}_e$
Muon	Muon	μ^-	105.7	-1	μ^+
	Muon neutrino ^b	ν_μ	$\approx 1 \times 10^{-7}$	0	$\bar{\nu}_\mu$
Tau	Tau	τ^-	1777	-1	τ^+
	Tau neutrino ^b	ν_τ	$\approx 1 \times 10^{-7}$	0	$\bar{\nu}_\tau$

^aAll leptons have spin quantum numbers of $\frac{1}{2}$ and are thus fermions.

^bThe neutrino masses have not been well determined.

The Conservation of Lepton Number

According to experiment, particle interactions involving leptons obey a conservation law for a quantum number called the **lepton number** L . Each (normal) particle in Table 44-2 is assigned $L = +1$, and each antiparticle is assigned $L = -1$. All other particles, which are not leptons, are assigned $L = 0$. Also according to experiment,



In all particle interactions, the net lepton number *for each family* is separately conserved.

Thus, there are actually three lepton numbers L_e , L_μ , L_τ , and the net of *each* must remain unchanged during any particle interaction. This experimental fact is called the law of **conservation of lepton number**.

We can illustrate this law by reconsidering the antimuon decay process shown in Eq. 44-8, which we now write more fully as

$$\mu^+ \rightarrow e^+ + \nu_e + \bar{\nu}_\mu. \quad (44-13)$$

Consider this first in terms of the muon family of leptons. The μ^+ is an antiparticle (see Table 44-2) and thus has the muon lepton number $L_\mu = -1$. The two particles e^+ and ν_e do not belong to the muon family and thus have $L_\mu = 0$. This leaves $\bar{\nu}_\mu$ on the right which, being an antiparticle, also has the muon lepton number $L_\mu = -1$. Thus, both sides of Eq. 44-13 have the same net muon lepton number—namely, $L_\mu = -1$; if they did not, the μ^+ would not decay by this process.

No members of the electron family appear on the left in Eq. 44-13; so there the net electron lepton number must be $L_e = 0$. On the right side of Eq. 44-13, the positron, being an antiparticle (again see Table 44-2), has the electron lepton number $L_e = -1$. The electron neutrino ν_e , being a particle, has the electron number $L_e = +1$. Thus, the net electron lepton number for these two particles on the right in Eq. 44-13 is also zero; the electron lepton number is also conserved in the process.

Because no members of the tau family appear on either side of Eq. 44-13, we must have $L_\tau = 0$ on each side. Thus, each of the lepton quantum numbers L_μ , L_e , and L_τ remains unchanged during the decay process of Eq. 44-13, their constant values being -1 , 0 , and 0 , respectively. This example is but one illustration of the conservation of lepton number; this law holds for all particle interactions. However, note that the law is based on (countless) experimental observations. We do not know *why* the law must be obeyed; we only know that this conservation law is part of the way our universe works.



CHECKPOINT 1

- (a) The π^+ meson decays by the process $\pi^+ \rightarrow \mu^+ + \nu$. To what lepton family does the neutrino ν belong? (b) Is this neutrino a particle or an antiparticle? (c) What is its lepton number?

44-5 The Hadrons

We are now ready to consider hadrons (baryons and mesons), those particles whose interactions are governed by the strong force. We start by adding another conservation law to our list: conservation of baryon number.

To develop this conservation law, let us consider the proton decay process

$$p \rightarrow e^+ + \nu_e. \quad (44-14)$$

This process *never* happens. We should be glad that it does not because otherwise all protons in the universe would gradually change into positrons, with disastrous consequences for us. Yet this decay process does not violate the conservation laws involving energy, linear momentum, or lepton number.

We account for the apparent stability of the proton—and for the absence of many other processes that might otherwise occur—by introducing a new quan-

tum number, the **baryon number** B , and a new conservation law, the **conservation of baryon number**:



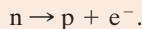
To every baryon we assign $B = +1$. To every antibaryon we assign $B = -1$. To all particles of other types we assign $B = 0$. A particle process cannot occur if it changes the net baryon number.

In the process of Eq. 44-14, the proton has a baryon number of $B = +1$ and the positron and neutrino both have a baryon number of $B = 0$. Thus, the process does not conserve baryon number and cannot occur.



CHECKPOINT 2

This mode of decay for a neutron is *not* observed:

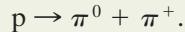


Which of the following conservation laws does this process violate: (a) energy, (b) angular momentum, (c) linear momentum, (d) charge, (e) lepton number, (f) baryon number? The masses are $m_n = 939.6 \text{ MeV}/c^2$, $m_p = 938.3 \text{ MeV}/c^2$, and $m_e = 0.511 \text{ MeV}/c^2$.

Sample Problem

Proton decay: conservation of quantum numbers, energy, and momentum

Determine whether a stationary proton can decay according to the scheme



Properties of the proton and the π^+ pion are listed in Table 44-1. The π^0 pion has zero charge, zero spin, and a mass energy of 135.0 MeV.

KEY IDEA

We need to see whether the proposed decay violates any of the conservation laws we have discussed.

Electric charge: We see that the net charge quantum number is initially +1 and finally $0 + 1$, or +1. Thus, charge is conserved by the decay. Lepton number is also conserved, because none of the three particles is a lepton and thus each lepton number is zero.

Linear momentum: Because the proton is stationary, with zero linear momentum, the two pions must merely move in opposite directions with equal magnitudes of linear momentum (so that their total linear momentum is also zero) to conserve linear momentum. The fact that linear momentum *can* be conserved means that the process does not violate the conservation of linear momentum.

Energy: Is there energy for the decay? Because the proton is stationary, that question amounts to asking whether the proton's mass energy is sufficient to produce the mass

energies and kinetic energies of the pions. To answer, we evaluate the Q of the decay:

$$\begin{aligned} Q &= \left(\frac{\text{initial total}}{\text{mass energy}} \right) - \left(\frac{\text{final total}}{\text{mass energy}} \right) \\ &= m_p c^2 - (m_0 c^2 + m_+ c^2) \\ &= 938.3 \text{ MeV} - (135.0 \text{ MeV} + 139.6 \text{ MeV}) \\ &= 663.7 \text{ MeV}. \end{aligned}$$

The fact that Q is positive indicates that the initial mass energy exceeds the final mass energy. Thus, the proton *does* have enough mass energy to create the pair of pions.

Spin: Is spin angular momentum conserved by the decay? This amounts to determining whether the net component S_z of spin angular momentum along some arbitrary z axis can be conserved by the decay. The spin quantum numbers s of the particles in the process are $\frac{1}{2}$ for the proton and 0 for both pions. Thus, for the proton the component S_z can be either $+\frac{1}{2}\hbar$ or $-\frac{1}{2}\hbar$ and for each pion it is $0\hbar$. We see that there is no way that S_z can be conserved. Hence, spin angular momentum is not conserved, and the proposed decay of the proton cannot occur.

Baryon number: The decay also violates the conservation of baryon number: The proton has a baryon number of $B = +1$, and both pions have a baryon number of $B = 0$. Thus, nonconservation of baryon number is another reason the proposed decay cannot occur.



Additional examples, video, and practice available at WileyPLUS

Sample Problem**Xi-minus decay: conservation of quantum numbers**

A particle called xi-minus and having the symbol Ξ^- decays as follows:

$$\Xi^- \rightarrow \Lambda^0 + \pi^-.$$

The Λ^0 particle (called lambda-zero) and the π^- particle are both unstable. The following decay processes occur in *cascade* until only relatively stable products remain:

$$\begin{aligned} \Lambda^0 &\rightarrow p + \pi^- & \pi^- &\rightarrow \mu^- + \bar{\nu}_\mu \\ \mu^- &\rightarrow e^- + \nu_\mu + \bar{\nu}_e \end{aligned}$$

- (a) Is the Ξ^- particle a lepton or a hadron? If the latter, is it a baryon or a meson?

KEY IDEAS

- (1) Only three families of leptons exist (Table 44-2) and none include the Ξ^- particle. Thus, the Ξ^- must be a hadron.
- (2) To answer the second question we need to determine the baryon number of the Ξ^- particle. If it is $+1$ or -1 , then the Ξ^- is a baryon. If, instead, it is 0 , then the Ξ^- is a meson.

Baryon number: To see, let us write the overall decay scheme, from the initial Ξ^- to the final relatively stable products, as

$$\Xi^- \rightarrow p + 2(e^- + \bar{\nu}_e) + 2(\nu_\mu + \bar{\nu}_\mu). \quad (44-15)$$

On the right side, the proton has a baryon number of $+1$ and each electron and neutrino has a baryon number of 0 . Thus, the net baryon number of the right side is $+1$. That must then be the baryon number of the lone Ξ^- particle on the left side. We conclude that the Ξ^- particle is a baryon.

- (b) Does the decay process conserve the three lepton numbers?

KEY IDEA

Any process must separately conserve the net lepton number for each lepton family of Table 44-2.

Lepton number: Let us first consider the electron lepton number L_e , which is $+1$ for the electron e^- , -1 for the anti-electron neutrino $\bar{\nu}_e$, and 0 for the other particles in the overall decay of Eq. 44-15. We see that the net L_e is 0 before the decay and $2[+1 + (-1)] + 2(0 + 0) = 0$ after the decay. Thus, the net electron lepton number is conserved. You can similarly show that the net muon lepton number and the net tau lepton number are also conserved.

- (c) What can you say about the spin of the Ξ^- particle?

KEY IDEA

The overall decay scheme of Eq. 44-15 must conserve the net spin component S_z .

Spin: We can determine the spin component S_z of the Ξ^- particle on the left side of Eq. 44-15 by considering the S_z components of the nine particles on the right side. All nine of those particles are spin- $\frac{1}{2}$ particles and thus can have S_z of either $+\frac{1}{2}\hbar$ or $-\frac{1}{2}\hbar$. No matter how we choose between those two possible values of S_z , the net S_z for those nine particles must be a *half-integer* times \hbar . Thus, the Ξ^- particle must have S_z of a *half-integer* times \hbar , and that means that its spin quantum number s must be a half-integer. (It is $\frac{1}{2}$.)



Additional examples, video, and practice available at WileyPLUS

44-6 Still Another Conservation Law

Particles have intrinsic properties in addition to the ones we have listed so far: mass, charge, spin, lepton number, and baryon number. The first of these additional properties was discovered when researchers observed that certain new particles, such as the kaon (K) and the sigma (Σ), always seemed to be produced in pairs. It seemed impossible to produce only one of them at a time. Thus, if a beam of energetic pions interacts with the protons in a bubble chamber, the reaction



often occurs. The reaction



which violates no conservation law known in the early days of particle physics, never occurs.

It was eventually proposed (by Murray Gell-Mann in the United States and independently by K. Nishijima in Japan) that certain particles possess a new

44-7 THE EIGHTFOLD WAY

1231

property, called **strangeness**, with its own quantum number S and its own conservation law. (Be careful not to confuse the symbol S here with spin.) The name *strangeness* arises from the fact that, before the identities of these particles were pinned down, they were known as “strange particles,” and the label stuck.

The proton, neutron, and pion have $S = 0$; that is, they are not “strange.” It was proposed, however, that the K^+ particle has strangeness $S = +1$ and that Σ^+ has $S = -1$. In the reaction of Eq. 44-16, the net strangeness is initially zero and finally zero; thus, the reaction conserves strangeness. However, in the reaction shown in Eq. 44-17, the final net strangeness is -1 ; thus, that reaction does not conserve strangeness and cannot occur. Apparently, then, we must add one more conservation law to our list—the **conservation of strangeness**:



Strangeness is conserved in interactions involving the strong force.

It may seem heavy-handed to invent a new property of particles just to account for a little puzzle like that posed by Eqs. 44-16 and 44-17. However, strangeness soon solved many other puzzles. Still, do not be misled by the whimsical character of the name. Strangeness is no more mysterious a property of particles than is charge. Both are properties that particles may (or may not) have; each is described by an appropriate quantum number. Each obeys a conservation law. Still other properties of particles have been discovered and given even more whimsical names, such as *charm* and *bottomness*, but all are perfectly legitimate properties. Let us see, as an example, how the new property of strangeness “earns its keep” by leading us to uncover important regularities in the properties of the particles.

44-7 The Eightfold Way

There are eight baryons—the neutron and the proton among them—that have a spin quantum number of $\frac{1}{2}$. Table 44-3 shows some of their other properties. Figure 44-3a shows the fascinating pattern that emerges if we plot the strangeness of these baryons against their charge quantum number, using a sloping axis for the charge quantum numbers. Six of the eight form a hexagon with the two remaining baryons at its center.

Let us turn now from the hadrons called baryons to the hadrons called mesons. Nine with a spin of zero are listed in Table 44-4. If we plot them on a sloping

Table 44-3
Eight Spin- $\frac{1}{2}$ Baryons

Particle	Symbol	Mass (MeV/c ²)	Quantum Numbers	
			Charge q	Strangeness S
Proton	p	938.3	+1	0
Neutron	n	939.6	0	0
Lambda	Λ^0	1115.6	0	-1
Sigma	Σ^+	1189.4	+1	-1
Sigma	Σ^0	1192.5	0	-1
Sigma	Σ^-	1197.3	-1	-1
Xi	Ξ^0	1314.9	0	-2
Xi	Ξ^-	1321.3	-1	-2

Table 44-4
Nine Spin-Zero Mesons^a

Particle	Symbol	Mass (MeV/c ²)	Quantum Numbers	
			Charge q	Strangeness S
Pion	π^0	135.0	0	0
Pion	π^+	139.6	+1	0
Pion	π^-	139.6	-1	0
Kaon	K^+	493.7	+1	+1
Kaon	K^-	493.7	-1	-1
Kaon	K^0	497.7	0	+1
Kaon	\bar{K}^0	497.7	0	-1
Eta	η	547.5	0	0
Eta prime	η'	957.8	0	0

^aAll mesons are bosons, having spins of 0, 1, 2, The ones listed here all have a spin of 0.

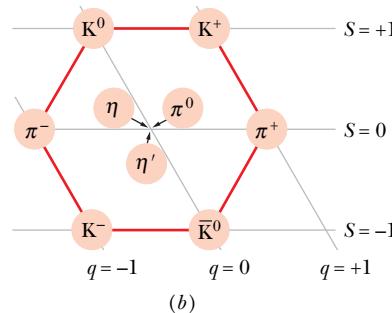
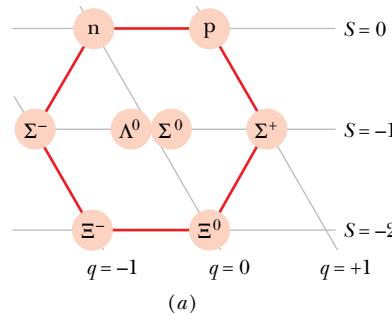


Fig. 44-3 (a) The eightfold way pattern for the eight spin- $\frac{1}{2}$ baryons listed in Table 44-3. The particles are represented as disks on a strangeness-charge plot, using a sloping axis for the charge quantum number. (b) A similar pattern for the nine spin-zero mesons listed in Table 44-4.

strangeness–charge diagram, as in Fig. 44-3b, the same fascinating pattern emerges! These and related plots, called the **eightfold way** patterns,* were proposed independently in 1961 by Murray Gell-Mann at the California Institute of Technology and by Yuval Ne’eman at Imperial College, London. The two patterns of Fig. 44-3 are representative of a larger number of symmetrical patterns in which groups of baryons and mesons can be displayed.

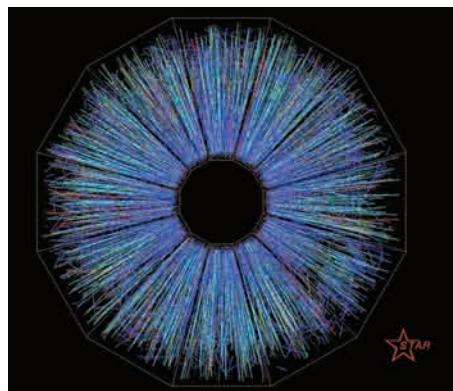
The symmetry of the eightfold way pattern for the spin- $\frac{3}{2}$ baryons (not shown here) calls for ten particles arranged in a pattern like that of the tenpins in a bowling alley. However, when the pattern was first proposed, only nine such particles were known; the “headpin” was missing. In 1962, guided by theory and the symmetry of the pattern, Gell-Mann made a prediction in which he essentially said:

There exists a spin- $\frac{3}{2}$ baryon with a charge of -1, a strangeness of -3, and a rest energy of about 1680 MeV. If you look for this omega minus particle (as I propose to call it), I think you will find it.

A team of physicists headed by Nicholas Samios of the Brookhaven National Laboratory took up the challenge and found the “missing” particle, confirming all its predicted properties. Nothing beats prompt experimental confirmation for building confidence in a theory!

The eightfold way patterns bear the same relationship to particle physics that the periodic table does to chemistry. In each case, there is a pattern of organization in which vacancies (missing particles or missing elements) stick out like sore thumbs, guiding experimenters in their searches. In the case of the periodic table, its very existence strongly suggests that the atoms of the elements are not fundamental particles but have an underlying structure. Similarly, the eightfold way patterns strongly suggest that the mesons and the baryons must have an underlying structure, in terms of which their properties can be understood. That structure can be explained in terms of the *quark model*, which we now discuss.

44-8 The Quark Model



The violent head-on collision of two 30 GeV beams of gold atoms in the RHIC accelerator at the Brookhaven National Laboratory. In the moment of collision, a gas of individual quarks and gluons was created. (Courtesy Brookhaven National Laboratory)

In 1964 Gell-Mann and George Zweig independently pointed out that the eightfold way patterns can be understood in a simple way if the mesons and the baryons are built up out of subunits that Gell-Mann called **quarks**. We deal first with three of them, called the *up quark* (symbol u), the *down quark* (symbol d), and the *strange quark* (symbol s), and we assign to them the properties displayed in Table 44-5. (The names of the quarks, along with those assigned to three other quarks that we shall meet later, have no meaning other than as convenient labels. Collectively, these names are called the *quark flavors*. We could just as well call them vanilla, chocolate, and strawberry instead of up, down, and strange.)

The fractional charge quantum numbers of the quarks may jar you a little. However, withhold judgment until you see how neatly these fractional charges account for the observed integer charges of the mesons and the baryons. In all normal situations, whether here on Earth or in an astronomical process, quarks are always bound up together in twos or threes for reasons that are still not well understood. Such requirements are our normal rule for quark combinations.

An exciting exception to the normal rule occurred in experiments at the RHIC particle collider at the Brookhaven National Laboratory. At the spot where two high-energy beams of gold nuclei collided head-on, the kinetic energy of the particles was so large that it matched the kinetic energy of

*The name is a borrowing from Eastern mysticism. The “eight” refers to the eight quantum numbers (only a few of which we have defined here) that are involved in the symmetry-based theory that predicts the existence of the patterns.

Table 44-5

The Quarks^a

Particle	Symbol	Mass (MeV/c ²)	Quantum Numbers			
			Charge <i>q</i>	Strangeness <i>S</i>	Baryon Number <i>B</i>	Antiparticle
Up	u	5	+ $\frac{2}{3}$	0	+ $\frac{1}{3}$	\bar{u}
Down	d	10	- $\frac{1}{3}$	0	+ $\frac{1}{3}$	\bar{d}
Charm	c	1500	+ $\frac{2}{3}$	0	+ $\frac{1}{3}$	\bar{c}
Strange	s	200	- $\frac{1}{3}$	-1	+ $\frac{1}{3}$	\bar{s}
Top	t	175 000	+ $\frac{2}{3}$	0	+ $\frac{1}{3}$	\bar{t}
Bottom	b	4300	- $\frac{1}{3}$	0	+ $\frac{1}{3}$	\bar{b}

^aAll quarks (including antiquarks) have spin $\frac{1}{2}$ and thus are fermions. The quantum numbers *q*, *S*, and *B* for each antiquark are the negatives of those for the corresponding quark.

particles that were present soon after the beginning of the universe (as we discuss in Section 44-14). The protons and neutrons of the gold nuclei were ripped apart to form a momentary gas of individual quarks. (The gas also contained gluons, the particles that normally hold quarks together, as we discuss in Section 44-9.) These experiments at RHIC may be the first time that quarks have been set free of one another since the universe began.

Quarks and Baryons

Each baryon is a combination of three quarks; some of the combinations are given in Fig. 44-4a. With regard to baryon number, we see that any three quarks (each with *B* = $+\frac{1}{3}$) yield a proper baryon (with *B* = +1).

Charges also work out, as we can see from three examples. The proton has a quark composition of uud, and so its charge quantum number is

$$q(uud) = \frac{2}{3} + \frac{2}{3} + \left(-\frac{1}{3}\right) = +1.$$

The neutron has a quark composition of udd, and its charge quantum number is therefore

$$q(udd) = \frac{2}{3} + \left(-\frac{1}{3}\right) + \left(-\frac{1}{3}\right) = 0.$$

The Σ^- (sigma-minus) particle has a quark composition of dds, and its charge quantum number is therefore

$$q(dds) = -\frac{1}{3} + \left(-\frac{1}{3}\right) + \left(-\frac{1}{3}\right) = -1.$$

The strangeness quantum numbers work out as well. You can check this by using Table 44-3 for the Σ^- strangeness number and Table 44-5 for the strangeness numbers of the dds quarks.

Note, however, that the mass of a proton, neutron, Σ^- , or any other baryon is *not* the sum of the masses of the constituent quarks. For example, the total mass of the three quarks in a proton is only 20 MeV/ c^2 , woefully less than the proton's mass of 938.3 MeV/ c^2 . Nearly all of the proton's mass is due to the internal energies of (1) the quark motion and (2) the fields that bind the quarks together (as discussed in Section 44-9). (Recall that mass is related to energy via Einstein's equation, which we can write as $m = E/c^2$.) Thus, because most of your mass is due to the protons and neutrons in your body, your mass (and therefore your weight on a bathroom scale) is primarily a measure of the energies of the quark motion and the quark-binding fields within you.

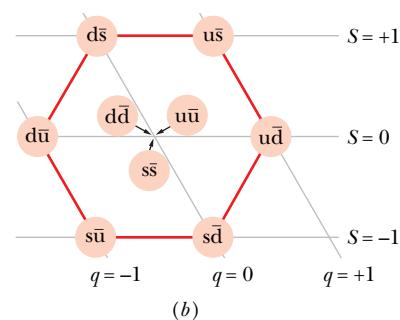
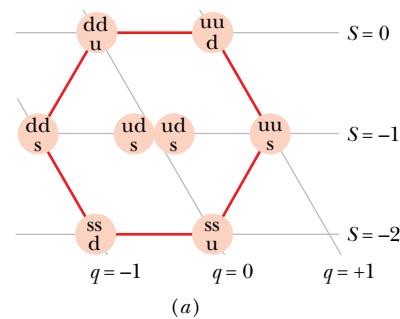


Fig. 44-4 (a) The quark compositions of the eight spin- $\frac{1}{2}$ baryons plotted in Fig. 44-3a. (Although the two central baryons share the same quark structure, they are different particles. The sigma is an excited state of the lambda, decaying into the lambda by emission of a gamma-ray photon.) (b) The quark compositions of the nine spin-zero mesons plotted in Fig. 44-3b.

Quarks and Mesons

Mesons are quark–antiquark pairs; some of their compositions are given in Fig. 44-4b. The quark–antiquark model is consistent with the fact that mesons are not baryons; that is, mesons have a baryon number $B = 0$. The baryon number for a quark is $+\frac{1}{3}$ and for an antiquark is $-\frac{1}{3}$; thus, the combination of baryon numbers in a meson is zero.

Consider the meson π^+ , which consists of an up quark u and an antidown quark \bar{d} . We see from Table 44-5 that the charge quantum number of the up quark is $+\frac{2}{3}$ and that of the antidown quark is $+\frac{1}{3}$ (the sign is opposite that of the down quark). This adds nicely to a charge quantum number of +1 for the π^+ meson; that is,

$$q(u\bar{d}) = \frac{2}{3} + \frac{1}{3} = +1.$$

All the charge and strangeness quantum numbers of Fig. 44-4b agree with those of Table 44-4 and Fig. 44-3b. Convince yourself that all possible up, down, and strange quark–antiquark combinations are used. Everything fits.



CHECKPOINT 3

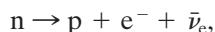
Is a combination of a down quark (d) and an antiup quark (\bar{u}) called (a) a π^0 meson, (b) a proton, (c) a π^- meson, (d) a π^+ meson, or (e) a neutron?

A New Look at Beta Decay

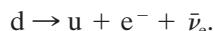
Let us see how beta decay appears from the quark point of view. In Eq. 42-24, we presented a typical example of this process:



After the neutron was discovered and Fermi had worked out his theory of beta decay, physicists came to view the fundamental beta-decay process as the changing of a neutron into a proton inside the nucleus, according to the scheme



in which the neutrino is identified more completely. Today we look deeper and see that a neutron (udd) can change into a proton (uud) by changing a down quark into an up quark. We now view the fundamental beta-decay process as



Thus, as we come to know more and more about the fundamental nature of matter, we can examine familiar processes at deeper and deeper levels. We see too that the quark model not only helps us to understand the structure of particles but also clarifies their interactions.

Still More Quarks

There are other particles and other eightfold way patterns that we have not discussed. To account for them, it turns out that we need to postulate three more quarks, the *charm quark* c, the *top quark* t, and the *bottom quark* b. Thus, a total of six quarks exist, as listed in Table 44-5.

Note that three quarks are exceptionally massive, the most massive of them (top) being almost 190 times more massive than a proton. To generate particles that contain such quarks, with such large mass energies, we must go to higher and higher energies, which is the reason that these three quarks were not discovered earlier.

The first particle containing a charm quark to be observed was the J/Ψ meson, whose quark structure is $c\bar{c}$. It was discovered simultaneously and independently in 1974 by groups headed by Samuel Ting at the Brookhaven National Laboratory and Burton Richter at Stanford University.

The top quark defied all efforts to generate it in the laboratory until 1995, when its existence was finally demonstrated in the Tevatron, a large particle accelerator at Fermilab. In this accelerator, protons and antiprotons, each with an energy of 0.9 TeV ($= 9 \times 10^{11}$ eV), are made to collide at the centers of two large particle detectors. In a very few cases, the colliding particles generate

a top–antitop ($t\bar{t}$) quark pair, which *very* quickly decays into particles that can be detected and thus can be used to infer the existence of the top–antitop pair.

Look back for a moment at Table 44-5 (the quark family) and Table 44-2 (the lepton family) and notice the neat symmetry of these two “six-packs” of particles, each dividing naturally into three corresponding two-particle families. In terms of what we know today, the quarks and the leptons seem to be truly fundamental particles having no internal structure.

Sample Problem

Quark composition of a xi-minus particle

The Ξ^- (xi-minus) particle is a baryon with a spin quantum number s of $\frac{1}{2}$, a charge quantum number q of -1 , and a strangeness quantum number S of -2 . Also, it does not contain a bottom quark. What combination of quarks makes up Ξ^- ?

Reasoning: Because the Ξ^- is a baryon, it must consist of three quarks (not two as for a meson).

Let us next consider the strangeness $S = -2$ of the Ξ^- . Only the strange quark s and the antistrange quark \bar{s} have nonzero values of strangeness (see Table 44-5). Further, because only the strange quark s has a *negative* value of strangeness, Ξ^- must contain that quark. In fact, for Ξ^- to have a strangeness of -2 , it must contain two strange quarks.

To determine the third quark, call it x , we can consider the other known properties of Ξ^- . Its charge quantum

number q is -1 , and the charge quantum number q of each strange quark is $-\frac{1}{3}$. Thus, the third quark x must have a charge quantum number of $-\frac{1}{3}$, so that we can have

$$\begin{aligned} q(\Xi^-) &= q(ssx) \\ &= -\frac{1}{3} + (-\frac{1}{3}) + (-\frac{1}{3}) = -1. \end{aligned}$$

Besides the strange quark, the only quarks with $q = -\frac{1}{3}$ are the down quark d and bottom quark b . Because the problem statement ruled out a bottom quark, the third quark must be a down quark. This conclusion is also consistent with the baryon quantum numbers:

$$\begin{aligned} B(\Xi^-) &= B(ssd) \\ &= \frac{1}{3} + \frac{1}{3} + \frac{1}{3} = +1. \end{aligned}$$

Thus, the quark composition of the Ξ^- particle is ssd.



Additional examples, video, and practice available at WileyPLUS

44-9 The Basic Forces and Messenger Particles

We turn now from cataloging the particles to considering the forces between them.

The Electromagnetic Force

At the atomic level, we say that two electrons exert electromagnetic forces on each other according to Coulomb’s law. At a deeper level, this interaction is described by a highly successful theory called **quantum electrodynamics** (QED). From this point of view, we say that each electron senses the presence of the other by exchanging photons with it.

We cannot detect these photons because they are emitted by one electron and absorbed by the other a very short time later. Because of their undetectable existence, we call them **virtual photons**. Because they communicate between the two interacting charged particles, we sometimes call these photons *messenger particles*.

If a stationary electron emits a photon and remains itself unchanged, energy is not conserved. The principle of conservation of energy is saved, however, by the uncertainty principle, written in the form

$$\Delta E \cdot \Delta t \approx \hbar. \quad (44-18)$$

Here we interpret this relation to mean that you can “overdraw” an amount of energy ΔE , violating conservation of energy, *provided* you “return” it within an

interval Δt given by $\hbar/\Delta E$ so that the violation cannot be detected. The virtual photons do just that. When, say, electron A emits a virtual photon, the overdraw in energy is quickly set right when that electron receives a virtual photon from electron B , and the violation is hidden by the inherent uncertainty.

The Weak Force

A theory of the weak force, which acts on all particles, was developed by analogy with the theory of the electromagnetic force. The messenger particles that transmit the weak force between particles, however, are not (massless) photons but massive particles, identified by the symbols W and Z . The theory was so successful that it revealed the electromagnetic force and the weak force as being different aspects of a single **electroweak force**. This accomplishment is a logical extension of the work of Maxwell, who revealed the electric and magnetic forces as being different aspects of a single *electromagnetic* force.

The electroweak theory was specific in predicting the properties of the messenger particles. Their charges and masses, for example, were predicted to be

Particle	Charge	Mass
W	$\pm e$	$80.4 \text{ GeV}/c^2$
Z	0	$91.2 \text{ GeV}/c^2$

Recall that the proton mass is only $0.938 \text{ GeV}/c^2$; these are massive particles! The 1979 Nobel Prize in physics was awarded to Sheldon Glashow, Steven Weinberg, and Abdus Salam for their development of the electroweak theory. The theory was confirmed in 1983 by Carlo Rubbia and his group at CERN, and the 1984 Nobel Prize in physics went to Rubbia and Simon van der Meer for this brilliant experimental work.

Some notion of the complexity of particle physics in this day and age can be found by looking at an earlier particle physics experiment that led to the Nobel Prize in physics—the discovery of the neutron. This vitally important discovery was a “tabletop” experiment, employing particles emitted by naturally occurring radioactive materials as projectiles; it was reported in 1932 under the title “Possible Existence of a Neutron,” the single author being James Chadwick.

The discovery of the W and Z messenger particles in 1983, by contrast, was carried out at a large particle accelerator, about 7 km in circumference and operating in the range of several hundred billion electron-volts. The principal particle detector alone weighed 20 MN. The experiment employed more than 130 physicists from 12 institutions in 8 countries, along with a large support staff.

The Strong Force

A theory of the strong force—that is, the force that acts between quarks to bind hadrons together—has also been developed. The messenger particles in this case are called **gluons** and, like the photon, they are predicted to be massless. The theory assumes that each “flavor” of quark comes in three varieties that, for convenience, have been labeled *red*, *yellow*, and *blue*. Thus, there are three up quarks, one of each color, and so on. The antiquarks also come in three colors, which we call *antired*, *antiyellow*, and *antiblue*. You must not think that quarks are actually colored, like tiny jelly beans. The names are labels of convenience, but (for once) they do have a certain formal justification, as you will see.

The force acting between quarks is called a **color force** and the underlying theory, by analogy with quantum electrodynamics (QED), is called **quantum chromodynamics** (QCD). Apparently, quarks can be assembled only in combinations that are *color-neutral*.

There are two ways to bring about color neutrality. In the theory of actual colors, red + yellow + blue yields white, which is color-neutral, and we use the same scheme in dealing with quarks. Thus we can assemble three quarks to form a baryon, provided one is a yellow quark, one is a red quark, and one is a blue quark. Antired + antiyellow + antiblue is also white, so that we can assemble three antiquarks (of the proper anticolors) to form an antibaryon. Finally, red + antired, or yellow + antiyellow, or blue + antiblue also yields white. Thus, we can assemble a quark–antiquark combination to form a meson. The color-neutral rule does not permit any other combination of quarks, and none are observed.

The color force not only acts to bind together quarks as baryons and mesons, but it also acts between such particles, in which case it has traditionally been called the strong force. Hence, not only does the color force bind together quarks to form protons and neutrons, but it also binds together the protons and neutrons to form nuclei.

Einstein's Dream

The unification of the fundamental forces of nature into a single force—which occupied Einstein's attention for much of his later life—is very much a current focus of research. We have seen that the weak force has been successfully combined with electromagnetism so that they may be jointly viewed as aspects of a single *electro-weak force*. Theories that attempt to add the strong force to this combination—called *grand unification theories* (GUTs)—are being pursued actively. Theories that seek to complete the job by adding gravity—sometimes called *theories of everything* (TOE)—are at an encouraging but speculative stage at this time.

44-10 A Pause for Reflection

Let us put what you have just learned in perspective. If all we are interested in is the structure of the world around us, we can get along nicely with the electron, the neutrino, the neutron, and the proton. As someone has said, we can operate “Spaceship Earth” quite well with just these particles. We can see a few of the more exotic particles by looking for them in the cosmic rays; however, to see most of them, we must build massive accelerators and look for them at great effort and expense.

The reason we must go to such effort is that—measured in energy terms—we live in a world of very low temperatures. Even at the center of the Sun, the value of kT is only about 1 keV. To produce the exotic particles, we must be able to accelerate protons or electrons to energies in the GeV and TeV range and higher.

Once upon a time the temperature everywhere *was* high enough to provide such energies. That time of extremely high temperatures occurred in the **big bang** beginning of the universe, when the universe (and both space and time) came into existence. Thus, one reason scientists study particles at high energies is to understand what the universe was like just after it began.

As we shall discuss shortly, *all* of space within the universe was initially tiny in extent, and the temperature of the particles within that space was incredibly high. With time, however, the universe expanded and cooled to lower temperatures, eventually to the size and temperature we see today.

Actually, the phrase “we see today” is complicated: When we look out into space, we are actually looking back in time because the light from the stars and galaxies has taken a long time to reach us. The most distant objects that we can detect are **quasars** (*quasistellar* objects), which are the extremely bright cores of galaxies that are as much as 13×10^9 ly from us. Each such core contains a gigantic black hole; as material (gas and even stars) is pulled into one of those black holes, the material heats up and radiates a tremendous amount of light, enough for us to detect in spite of the huge distance. We therefore “see” a quasar not as it looks today but rather as it once was, when that light began its journey to us billions of years ago.

44-11 The Universe Is Expanding

As we saw in Section 37-10, it is possible to measure the relative speeds at which galaxies are approaching us or receding from us by measuring the shifts in the wavelength of the light they emit. If we look only at distant galaxies, beyond our immediate galactic neighbors, we find an astonishing fact: They are *all* moving away (receding) from us! In 1929 Edwin P. Hubble connected the recession speed v of a galaxy and its distance r from us—they are directly proportional:

$$v = Hr \quad (\text{Hubble's law}), \quad (44-19)$$

in which H is called the **Hubble constant**. The value of H is usually measured in the unit kilometers per second-megaparsec ($\text{km/s} \cdot \text{Mpc}$), where the megaparsec is a length unit commonly used in astrophysics and astronomy:

$$1 \text{ Mpc} = 3.084 \times 10^{19} \text{ km} = 3.260 \times 10^6 \text{ ly}. \quad (44-20)$$

The Hubble constant H has not had the same value since the universe began. Determining its current value is extremely difficult because doing so involves measurements of very distant galaxies. However, from measurements that we shall discuss in Section 44-14, the Hubble constant is now known to be

$$H = 71.0 \text{ km/s} \cdot \text{Mpc} = 21.8 \text{ mm/s} \cdot \text{ly}. \quad (44-21)$$

We interpret the recession of the galaxies to mean that the universe is expanding, much as the raisins in what is to be a loaf of raisin bread grow farther apart as the dough expands. Observers on all other galaxies would find that distant galaxies were rushing away from them also, in accordance with Hubble's law. In keeping with our analogy, we can say that no raisin (galaxy) has a unique or preferred view.

Hubble's law is consistent with the hypothesis that the universe began with the big bang and has been expanding ever since. If we assume that the rate of expansion has been constant (that is, the value of H has been constant), then we can estimate the age T of the universe by using Eq. 44-19. Let us also assume that since the big bang, any given part of the universe (say, a galaxy) has been receding from our location at a speed v given by Eq. 44-19. Then the time required for the given part to recede a distance r is

$$T = \frac{r}{v} = \frac{r}{Hr} = \frac{1}{H} \quad (\text{estimated age of universe}). \quad (44-22)$$

For the value of H in Eq. 44-21, T works out to be 13.8×10^9 y. Much more sophisticated studies of the expansion of the universe put T at 13.7×10^9 y.

Sample Problem

Using Hubble's law to relate distance and recessional speed

The wavelength shift in the light from a particular quasar indicates that the quasar has a recessional speed of 2.8×10^8 m/s (which is 93% of the speed of light). Approximately how far from us is the quasar?

KEY IDEA

We assume that the distance and speed are related by Hubble's law.

Calculation: From Eqs. 44-19 and 44-21, we find

$$\begin{aligned} r &= \frac{v}{H} = \frac{2.8 \times 10^8 \text{ m/s}}{21.8 \text{ mm/s} \cdot \text{ly}} (1000 \text{ mm/m}) \\ &= 12.8 \times 10^9 \text{ ly}. \end{aligned} \quad (\text{Answer})$$

This is only an approximation because the quasar has not always been receding from our location at the same speed v ; that is, H has not had its current value throughout the time during which the universe has been expanding.



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Sample Problem**Using Hubble's law to relate distance and Doppler shift**

A particular emission line detected in the light from a galaxy has a detected wavelength $\lambda_{\text{det}} = 1.1\lambda$, where λ is the proper wavelength of the line. What is the galaxy's distance from us?

KEY IDEAS

- (1) We assume that Hubble's law ($v = Hr$) applies to the recession of the galaxy.
- (2) We also assume that the astronomical Doppler shift of Eq. 37-36 ($v = c \Delta\lambda/\lambda$, for $v \ll c$) applies to the shift in wavelength due to the recession.

Calculations: We can then set the right side of these two equations equal to each other to write

$$Hr = \frac{c \Delta\lambda}{\lambda}, \quad (44-23)$$

which leads us to

$$r = \frac{c \Delta\lambda}{H\lambda}. \quad (44-24)$$

In this equation,

$$\Delta\lambda = \lambda_{\text{det}} - \lambda = 1.1\lambda - \lambda = 0.1\lambda.$$

Substituting this into Eq. 44-24 then gives us

$$\begin{aligned} r &= \frac{c(0.1\lambda)}{H\lambda} = \frac{0.1c}{H} \\ &= \frac{(0.1)(3.0 \times 10^8 \text{ m/s})}{21.8 \text{ mm/s} \cdot \text{ly}} (1000 \text{ mm/m}) \\ &= 1.4 \times 10^9 \text{ ly.} \end{aligned} \quad (\text{Answer})$$



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44-12 The Cosmic Background Radiation

In 1965 Arno Penzias and Robert Wilson, of what was then the Bell Telephone Laboratories, were testing a sensitive microwave receiver used for communications research. They discovered a faint background "hiss" that remained unchanged in intensity no matter where their antenna was pointed. It soon became clear that Penzias and Wilson were observing a **cosmic background radiation**, generated in the early universe and filling all space almost uniformly. Currently this radiation has a maximum intensity at a wavelength of 1.1 mm, which lies in the microwave region of electromagnetic radiation (or light, for short). The wavelength distribution of this radiation matches the wavelength distribution of light that would be emitted by a laboratory enclosure with walls at a temperature of 2.7 K. Thus, for the cosmic background radiation, we say that the enclosure is the entire universe and that the universe is at an (average) temperature of 2.7 K. For their discovery of the cosmic background radiation, Penzias and Wilson were awarded the 1978 Nobel Prize in physics.

As we discuss in Section 44-14, the cosmic background radiation is now known to be light that has been in flight across the universe since shortly after the universe began billions of years ago. When the universe was even younger, light could scarcely go any significant distance without being scattered by all the individual, high-speed particles along its path. If a light ray started from, say, point *A*, it would be scattered in so many directions that if you could have intercepted part of it, you would have not been able to tell that it originated at point *A*. However, after the particles began to form atoms, the scattering of light greatly decreased. A light ray from point *A* might then be able to travel for billions of years without being scattered. This light is the cosmic background radiation.

As soon as the nature of the radiation was recognized, researchers wondered, "Can we use this incoming radiation to distinguish the points at which it originated, so that we then can produce an image of the early universe, back when atoms first formed and light scattering largely ceased?" The answer is yes, and that image is coming up in Section 44-14.

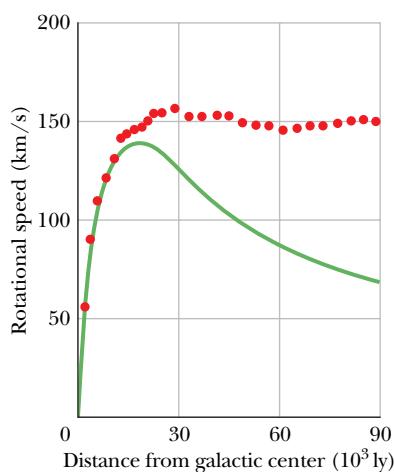


Fig. 44-5 The rotational speed of stars in a typical galaxy as a function of their distance from the galactic center. The theoretical solid curve shows that if a galaxy contained only the mass that is visible, the observed rotational speed would drop off with distance at large distances. The dots are the experimental data, which show that the rotational speed is approximately constant at large distances.

44-13 Dark Matter

At the Kitt Peak National Observatory in Arizona, Vera Rubin and her co-worker Kent Ford measured the rotational rates of a number of distant galaxies. They did so by measuring the Doppler shifts of bright clusters of stars located within each galaxy at various distances from the galactic center. As Fig. 44-5 shows, their results were surprising: The orbital speed of stars at the outer visible edge of the galaxy is about the same as that of stars close to the galactic center.

As the solid curve in Fig. 44-5 attests, that is not what we would expect to find if all the mass of the galaxy were represented by visible light. Nor is the pattern found by Rubin and Ford what we find in the solar system. For example, the orbital speed of Pluto (the planet most distant from the Sun) is only about one-tenth that of Mercury (the planet closest to the Sun).

The only explanation for the findings of Rubin and Ford that is consistent with Newtonian mechanics is that a typical galaxy contains much more matter than what we can actually see. In fact, the visible portion of a galaxy represents only about 5 to 10% of the total mass of the galaxy. In addition to these studies of galactic rotation, many other observations lead to the conclusion that the universe abounds in matter that we cannot see. This unseen matter is called **dark matter** because either it does not emit light or its light emission is too dim for us to detect.

Normal matter (such as stars, planets, dust, and molecules) is often called **baryonic matter** because its mass is primarily due to the combined mass of the protons and neutrons (baryons) it contains. (The mass of the electrons is neglected because the mass of an electron is so small relative to the mass of a proton or a neutron.) Some of the normal matter, such as burned-out stars and dim interstellar gas, is part of the dark matter in a galaxy.

However, according to various calculations, this dark normal matter is only a small part of the total dark matter. The rest is called **nonbaryonic dark matter** because it does not contain protons and neutrons. We know of only one member of this type of dark matter—the neutrinos. Although the mass of a neutrino is very small relative to the mass of a proton or neutron, the number of neutrinos in a galaxy is huge and thus the total mass of the neutrinos is large. Nevertheless, calculations indicate that not even the total mass of the neutrinos is enough to account for the total mass of the nonbaryonic dark matter. In spite of over a hundred years in which elementary particles have been detected and studied, the particles that make up the rest of this type of dark matter are undetected and their nature is unknown. Because we have no experience with them, they must interact only gravitationally with the common particles.

44-14 The Big Bang

In 1985, a physicist remarked at a scientific meeting:

It is as certain that the universe started with a big bang about 15 billion years ago as it is that the Earth goes around the Sun.

This strong statement suggests the level of confidence in which the big bang theory, first advanced by Belgian physicist Georges Lemaître, is held by those who study these matters. However, you must not imagine that the big bang was like the explosion of some gigantic firecracker and that, in principle at least, you could have stood to one side and watched. There was no “one side” because the big bang represents the beginning of spacetime itself. From the point of view of our present universe, there is no position in space to which you can point and say, “The big bang happened there.” It happened everywhere.

Moreover, there was no “before the big bang,” because time *began* with that creation event. In this context, the word “before” loses its meaning. We can,

however, conjecture about what went on during succeeding intervals of time after the big bang (Fig. 44-6).

$t \approx 10^{-43}$ s. This is the earliest time at which we can say anything meaningful about the development of the universe. It is at this moment that the concepts of space and time come to have their present meanings and the laws of physics as we know them become applicable. At this instant, the entire universe (that is, the *entire* spatial extent of the universe) is much smaller than a proton and its temperature is about 10^{32} K. Quantum fluctuations in the fabric of spacetime are the seeds that will eventually lead to the formation of galaxies, clusters of galaxies, and superclusters of galaxies.

$t \approx 10^{-34}$ s. By this moment the universe has undergone a tremendously rapid inflation, increasing in size by a factor of about 10^{30} , causing the formation of matter in a distribution set by the initial quantum fluctuations. The universe has become a hot soup of photons, quarks, and leptons at a temperature of about 10^{27} K, which is too hot for protons and neutrons to form.

$t \approx 10^{-4}$ s. Quarks can now combine to form protons and neutrons and their antiparticles. The universe has now cooled to such an extent by continued (but much slower) expansion that photons lack the energy needed to break up these new particles. Particles of matter and antimatter collide and annihilate each other. There is a slight excess of matter, which, failing to find annihilation partners, survives to form the world of matter that we know today.

$t \approx 1$ min. The universe has now cooled enough so that protons and neutrons, in colliding, can stick together to form the low-mass nuclei ^2H , ^3He , ^4He , and ^7Li . The predicted relative abundances of these nuclides are just what we observe in the universe today. Also, there is plenty of radiation present at $t \approx 1$ min, but this light cannot travel far before it interacts with a nucleus. Thus the universe is opaque.

$t \approx 379\,000$ y. The temperature has now fallen to 2970 K, and electrons can stick to bare nuclei when the two collide, forming atoms. Because light does not interact appreciably with (uncharged) particles, such as neutral atoms, the light is now free to travel great distances. This radiation forms the cosmic background radiation discussed in Section 44-12. Atoms of hydrogen and helium, under the influence of gravity, begin to clump together, eventually starting the formation of galaxies and stars, but until then, the universe is relatively dark (Fig. 44-6).

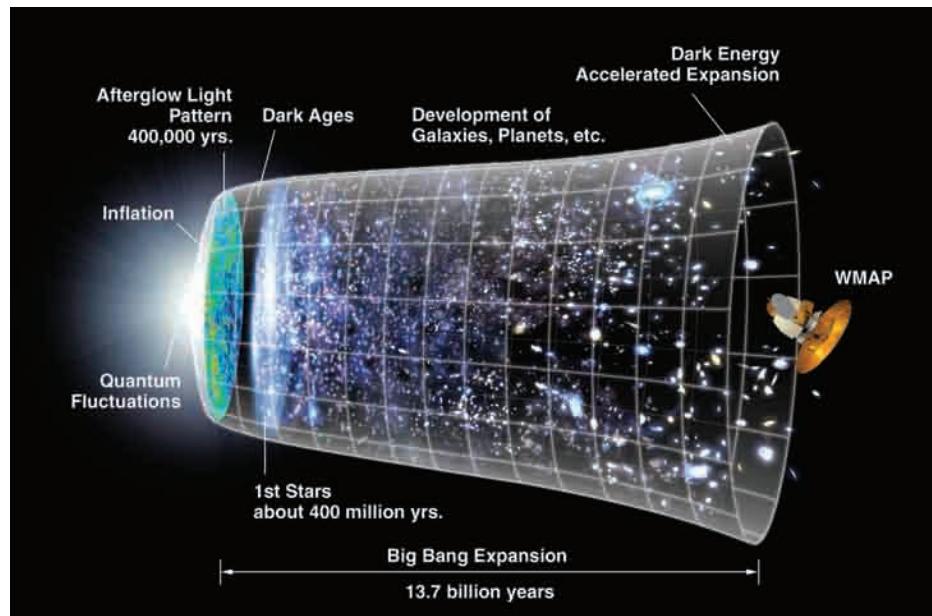


Fig. 44-6 An illustration of the universe from the initial quantum fluctuations just after $t = 0$ (at the left) to the current accelerated expansion, 13.7×10^9 y later (at the right). Don't take the illustration literally—there is *no* such “external view” of the universe because there is *no* exterior to the universe. (Courtesy NASA)

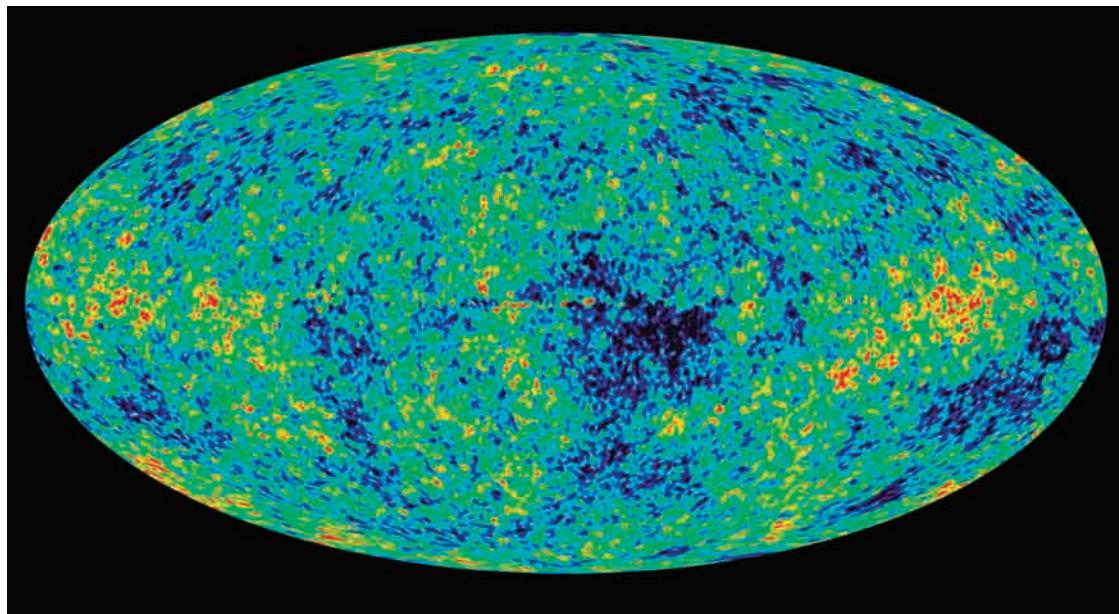


Fig. 44-7 This color-coded image is effectively a photograph of the universe when it was only 379 000 y old, which was about 13.7×10^9 y ago. This is what you would have seen then as you looked away in all directions (the view has been condensed to this oval). Patches of light from collections of atoms stretch across the “sky,” but galaxies, stars, and planets have not yet formed. (Courtesy WMAP Science Team/NASA)

Early measurements suggested that the cosmic background radiation is uniform in all directions, implying that 379 000 y after the big bang all matter in the universe was uniformly distributed. This finding was most puzzling because matter in the present universe is not uniformly distributed, but instead is collected in galaxies, clusters of galaxies, and superclusters of galactic clusters. There are also vast *voids* in which there is relatively little matter, and there are regions so crowded with matter that they are called *walls*. If the big bang theory of the beginning of the universe is even approximately correct, the seeds for this non-uniform distribution of matter must have been in place before the universe was 379 000 y old and now should show up as a nonuniform distribution of the microwave background radiation.

In 1992, measurements made by NASA’s Cosmic Background Explorer (COBE) satellite revealed that the background radiation is, in fact, not perfectly uniform. In 2003, measurements by NASA’s Wilkinson Microwave Anisotropy Probe (WMAP) greatly increased our resolution of this nonuniformity. The resulting image (Fig. 44-7) is effectively a color-coded photograph of the universe when it was only 379 000 y old. As you can see from the variations in the colors, large-scale collecting of matter had already begun. Thus, the big bang theory and the theory of inflation at $t \approx 10^{-34}$ s are on the right track.

The Accelerated Expansion of the Universe

Recall from Section 13-9 the statement that mass causes curvature of space. Now that we have seen that mass is a form of energy, as given by Einstein’s equation $E = mc^2$, we can generalize the statement: energy can cause curvature of space. This certainly happens to the space around the energy packed into a black hole and, more weakly, to the space around any other astronomical body, but is the space of the universe as a whole curved by the energy the universe contains?

The question was answered first by the 1992 COBE measurements of the cosmic background radiation. It was then answered more definitively by the 2003 WMAP measurements that produced the image in Fig. 44-7. The spots we see in that image are the original sources of the cosmic background radiation, and the angular distribution of the spots reveals the curvature of the universe through which the light has to travel to reach us. If adjacent spots subtend either more than 1° (Fig. 44-8a) or less than 1° (Fig. 44-8b) in the detector’s view (or our view) into the uni-

verse, then the universe is curved. Analysis of the spot distribution in the WMAP image shows that the spots subtend about 1° (Fig. 44-8c), which means that the universe is *flat* (having no curvature). Thus, the initial curvature the universe presumably had when it began must have been flattened out by the rapid inflation the universe underwent at $t \approx 10^{-34}$ s.

This flatness poses a very difficult problem for physicists because it requires that the universe contain a certain amount of energy (as mass or otherwise). The trouble is that all estimations of the amount of energy in the universe (both in known forms and in the form of the unknown type of dark matter) fall dramatically short of the required amount. Indeed, about two-thirds of the required energy is missing from the estimations.

One theory proposed about this missing energy gave it the gothic name of *dark energy* and predicted that it has the strange property of causing the expansion of the universe to accelerate. Until 1998, determining whether the expansion is, in fact, accelerating was very difficult because it requires measuring distances to very distant astronomical bodies where the acceleration might show up.

In 1998, however, advances in astronomical technology allowed astronomers to detect a certain type of supernovae at very great distances. More important, the astronomers could measure the duration of the burst of light from such a supernova. The duration reveals the brightness of the supernova that would be seen by an observer near the supernova. By measuring the brightness of the supernova as seen from Earth, astronomers could then determine the distance to the supernova. From the redshift of the light from the galaxy containing the supernova, astronomers could also determine how fast the galaxy is receding from us. Combining all this information, they could then calculate the expansion rate of the universe. The conclusion is that the expansion is indeed accelerating as predicted by the theory of dark energy (Fig. 44-6). However, we still scarcely have a clue as to what this dark energy is.

44-15 A Summing Up

Let us, in these closing paragraphs, consider where our rapidly accumulating store of knowledge about the universe is leading us. That it provides satisfaction to a host of curiosity-motivated physicists and astronomers is beyond dispute. However, some view it as a humbling experience in that each increase in knowledge seems to reveal more clearly our own relative insignificance in the grand scheme of things. Thus, in roughly chronological order, we humans have come to realize that

Our Earth is not the center of the solar system.

Our Sun is but one star among many in our galaxy.

Our galaxy is but one of many, and our Sun is an insignificant star in it.

Our Earth has existed for perhaps only a third of the age of the universe and will surely disappear when our Sun burns up its fuel and becomes a red giant.

Our species has inhabited Earth for less than a million years—a blink in cosmological time.

Although our position in the universe may be insignificant, the laws of physics that we have discovered (uncovered?) seem to hold throughout the universe and—as far as we know—have held since the universe began and will continue to hold for all future time. At least, there is no evidence that other laws hold in other parts of the universe. Thus, until someone complains, we are entitled to stamp the laws of physics “Discovered on Earth.” Much remains to be discovered: “*The universe is full of magical things, patiently waiting for our wits to grow sharper.*” That declaration allows us to answer one last time the question “What is physics?” that we have explored repeatedly in this book. Physics is the gateway to those magical things.

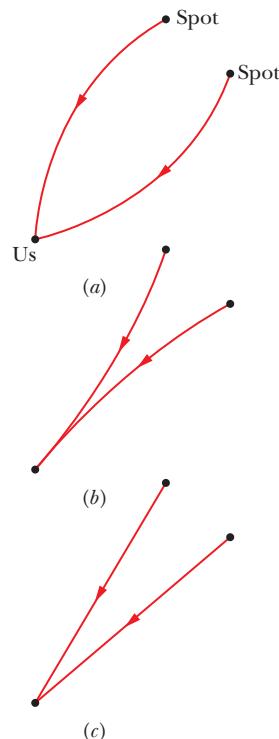


Fig. 44-8 Light rays from two adjacent spots in our view of the cosmic background radiation would reach us at an angle (a) greater than 1° or (b) less than 1° if the space along the light-ray paths through the universe were curved. (c) An angle of 1° means that the space is not curved.

REVIEW & SUMMARY

Leptons and Quarks Current research supports the view that all matter is made of six kinds of **leptons** (Table 44-2), six kinds of **quarks** (Table 44-5), and 12 **antiparticles**, one corresponding to each lepton and each quark. All these particles have spin quantum numbers equal to $\frac{1}{2}$ and are thus **fermions** (particles with half-integer spin quantum numbers).

The Interactions Particles with electric charge interact through the electromagnetic force by exchanging **virtual photons**. Leptons can interact with each other and with quarks through the **weak force**, via massive W and Z particles as messengers. In addition, quarks interact with each other through the **color force**. The electromagnetic and weak forces are different manifestations of the same force, called the **electroweak force**.

Leptons Three of the leptons (the **electron**, **muon**, and **tau**) have electric charge equal to $-1e$. There are also three uncharged **neutrinos** (also leptons), one corresponding to each of the charged leptons. The antiparticles for the charged leptons have positive charge.

Quarks The six quarks (up, down, strange, charm, bottom, and top, in order of increasing mass) each have baryon number $\frac{1}{3}$ and charge equal to either $+\frac{2}{3}e$ or $-\frac{1}{3}e$. The strange quark has strange-

ness -1 , whereas the others all have strangeness 0 . These four algebraic signs are reversed for the antiquarks.

Hadrons: Baryons and Mesons Quarks combine into strongly interacting particles called **hadrons**. **Baryons** are hadrons with half-integer spin quantum numbers ($\frac{1}{2}$ or $\frac{3}{2}$). **Mesons** are hadrons with integer spin quantum numbers (0 or 1) and thus are **bosons**. Baryons are fermions. Mesons have baryon number equal to zero; baryons have baryon number equal to $+1$ or -1 . **Quantum chromodynamics** predicts that the possible combinations of quarks are either a quark with an antiquark, three quarks, or three antiquarks (this prediction is consistent with experiment).

Expansion of the Universe Current evidence strongly suggests that the universe is expanding, with the distant galaxies moving away from us at a rate v given by **Hubble's law**:

$$v = Hr \quad (\text{Hubble's law}). \quad (44-19)$$

Here we take H , the **Hubble constant**, to have the value

$$H = 71.0 \text{ km/s} \cdot \text{Mpc} = 21.8 \text{ mm/s} \cdot \text{ly}. \quad (44-21)$$

The expansion described by Hubble's law and the presence of ubiquitous background microwave radiation reveal that the universe began in a "big bang" 13.7 billion years ago.

**** View All Solutions Here ******Q U E S T I O N S**

- 1 An electron cannot decay into two neutrinos. Which of the following conservation laws would be violated if it did: (a) energy, (b) angular momentum, (c) charge, (d) lepton number, (e) linear momentum, (f) baryon number?

- 2 Which of the eight pions in Fig. 44-2b has the least kinetic energy?

- 3 Figure 44-9 shows the paths of two particles circling in a uniform magnetic field. The particles have the same magnitude of charge but opposite signs. (a) Which path corresponds to the more massive particle? (b) If the magnetic field is directed into the plane of the page, is the more massive particle positively or negatively charged?

- 4 A proton has enough mass energy to decay into a shower made up of electrons, neutrinos, and their antiparticles. Which of the following conservation laws would necessarily be violated if it did: electron lepton number or baryon number?

- 5 A proton cannot decay into a neutron and a neutrino. Which of the following conservation laws would be violated if it did: (a) energy (assume the proton is stationary), (b) angular momentum, (c) charge, (d) lepton number, (e) linear momentum, (f) baryon number?

- 6 Does the proposed decay $\Lambda^0 \rightarrow p + K^-$ conserve (a) electric charge, (b) spin angular momentum, and (c) strangeness? (d) If the original particle is stationary, is there enough energy to create the decay products?

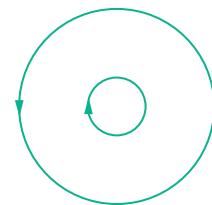


Fig. 44-9

Question 3.

- 7 Not only particles such as electrons and protons but also entire atoms can be classified as fermions or bosons, depending on whether their overall spin quantum numbers are, respectively, half-integral or integral. Consider the helium isotopes ^3He and ^4He . Which of the following statements is correct? (a) Both are fermions. (b) Both are bosons. (c) ^4He is a fermion, and ^3He is a boson. (d) ^3He is a fermion, and ^4He is a boson. (The two helium electrons form a closed shell and play no role in this determination.)

- 8 Three cosmologists have each plotted a line on the Hubble-like graph of Fig. 44-10. If we calculate the corresponding age of the universe from the three plots, rank the plots according to that age, greatest first.

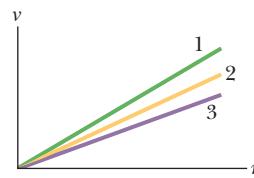


Fig. 44-10

Question 8.

- 9 A Σ^+ particle has these quantum numbers: strangeness $S = -1$, charge $q = +1$, and spin $s = \frac{1}{2}$. Which of the following quark combinations produces it: (a) dds, (b) ss, (c) uus, (d) ssu, or (e) uu \bar{s} ?

- 10 As we have seen, the π^- meson has the quark structure d \bar{u} . Which of the following conservation laws would be violated if a π^- were formed, instead, from a d quark and a u quark: (a) energy, (b) angular momentum, (c) charge, (d) lepton number, (e) linear momentum, (f) baryon number?

- 11 Consider the neutrino whose symbol is $\bar{\nu}_e$. (a) Is it a quark, a lepton, a meson, or a baryon? (b) Is it a particle or an antiparticle? (c) Is it a boson or a fermion? (d) Is it stable against spontaneous decay?

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PROBLEMS



Tutoring problem available (at instructor's discretion) in WileyPLUS and WebAssign



Worked-out solution available in Student Solutions Manual



Number of dots indicates level of problem difficulty

Additional information available in *The Flying Circus of Physics* and at flyingcircusofphysics.com

WWW Worked-out solution is at

ILW Interactive solution is at

<http://www.wiley.com/college/halliday>**sec. 44-3 An Interlude**

- 1 A positively charged pion decays by Eq. 44-7: $\pi^+ \rightarrow \mu^+ + \nu$. What must be the decay scheme of the negatively charged pion? (*Hint:* The π^- is the antiparticle of the π^+ .)

- 2 Certain theories predict that the proton is unstable, with a half-life of about 10^{32} years. Assuming that this is true, calculate the number of proton decays you would expect to occur in one year in the water of an Olympic-sized swimming pool holding 4.32×10^5 L of water.

- 3 An electron and a positron undergo pair annihilation (Eq. 44-5). If they had approximately zero kinetic energy before the annihilation, what is the wavelength of each γ produced by the annihilation?

- 4 A neutral pion initially at rest decays into two gamma rays: $\pi^0 \rightarrow \gamma + \gamma$. Calculate the wavelength of the gamma rays. Why must they have the same wavelength?

- 5 An electron and a positron are separated by distance r . Find the ratio of the gravitational force to the electric force between them. From the result, what can you conclude concerning the forces acting between particles detected in a bubble chamber? (Should gravitational interactions be considered?)

- 6 (a) A stationary particle 1 decays into particles 2 and 3, which move off with equal but oppositely directed momenta. Show that the kinetic energy K_2 of particle 2 is given by

$$K_2 = \frac{1}{2E_1} [(E_1 - E_2)^2 - E_3^2],$$

- where E_1 , E_2 , and E_3 are the rest energies of the particles. (b) A stationary positive pion π^+ (rest energy 139.6 MeV) can decay to an antimuon μ^+ (rest energy 105.7 MeV) and a neutrino ν (rest energy approximately 0). What is the resulting kinetic energy of the antimuon?

- 7 The rest energy of many short-lived particles cannot be measured directly but must be inferred from the measured momenta and known rest energies of the decay products. Consider the ρ^0 meson, which decays by the reaction $\rho^0 \rightarrow \pi^+ + \pi^-$. Calculate the rest energy of the ρ^0 meson given that the oppositely directed momenta of the created pions each have magnitude 358.3 MeV/c. See Table 44-4 for the rest energies of the pions.

- 8 A positive tau (τ^+ , rest energy = 1777 MeV) is moving with 2200 MeV of kinetic energy in a circular path perpendicular to a uniform 1.20 T magnetic field. (a) Calculate the momentum of the tau in kilogram-meters per second. Relativistic effects must be considered. (b) Find the radius of the circular path.

- 9 Observations of neutrinos emitted by the supernova SN1987a (Fig. 43-12b) place an upper limit of 20 eV on the rest energy of the electron neutrino. If the rest energy of the electron neutrino were, in fact, 20 eV, what would be the speed difference between light and a 1.5 MeV electron neutrino?

- 10 A neutral pion has a rest energy of 135 MeV and a

mean life of 8.3×10^{-17} s. If it is produced with an initial kinetic energy of 80 MeV and decays after one mean lifetime, what is the longest possible track this particle could leave in a bubble chamber? Use relativistic time dilation.

sec. 44-6 Still Another Conservation Law

- 11 Which conservation law is violated in each of these proposed decays? Assume that the initial particle is stationary and the decay products have zero orbital angular momentum. (a) $\mu^- \rightarrow e^- + \nu_\mu$; (b) $\mu^- \rightarrow e^+ + \nu_e + \bar{\nu}_\mu$; (c) $\mu^+ \rightarrow \pi^+ + \nu_\mu$.

- 12 The A_2^+ particle and its products decay according to the scheme

$$\begin{array}{ll} A_2^+ \rightarrow \rho^0 + \pi^+, & \mu^+ \rightarrow e^+ + \nu + \bar{\nu}, \\ \rho^0 \rightarrow \pi^+ + \pi^-, & \pi^- \rightarrow \mu^- + \bar{\nu}, \\ \mu^+ \rightarrow \mu^+ + \nu, & \mu^- \rightarrow e^- + \nu + \bar{\nu}. \end{array}$$

- (a) What are the final stable decay products? From the evidence, (b) is the A_2^+ particle a fermion or a boson and (c) is it a meson or a baryon? (d) What is its baryon number?

sec. 44-7 The Eightfold Way

- 13 Show that if, instead of plotting strangeness S versus charge q for the spin- $\frac{1}{2}$ baryons in Fig. 44-3a and for the spin-zero mesons in Fig. 44-3b, we plot the quantity $Y = B + S$ versus the quantity $T_z = q - \frac{1}{2}(B + S)$, we get the hexagonal patterns without using sloping axes. (The quantity Y is called *hypercharge*, and T_z is related to a quantity called *isospin*.)

- 14 Calculate the disintegration energy of the reactions (a) $\pi^+ + p \rightarrow \Sigma^+ + K^+$ and (b) $K^- + p \rightarrow \Lambda^0 + \pi^0$.

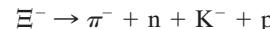
- 15 Which conservation law is violated in each of these proposed reactions and decays? (Assume that the products have zero orbital angular momentum.) (a) $\Lambda^0 \rightarrow p + K^-$; (b) $\Omega^- \rightarrow \Sigma^- + \pi^0$ ($S = -3, q = -1, m = 1672 \text{ MeV}/c^2$, and $m_s = \frac{3}{2}$ for Ω^-); (c) $K^- + p \rightarrow \Lambda^0 + \pi^+$.

- 16 Does the proposed reaction



- conserve (a) charge, (b) baryon number, (c) electron lepton number, (d) spin angular momentum, (e) strangeness, and (f) muon lepton number?

- 17 Does the proposed decay process



- conserve (a) charge, (b) baryon number, (c) spin angular momentum, and (d) strangeness?

- 18 By examining strangeness, determine which of the following decays or reactions proceed via the strong interaction: (a) $K^0 \rightarrow \pi^+ + \pi^-$; (b) $\Lambda^0 + p \rightarrow \Sigma^+ + n$; (c) $\Lambda^0 \rightarrow p + \pi^-$; (d) $K^- + p \rightarrow \Lambda^0 + \pi^0$.

- 19 The reaction $\pi^+ + p \rightarrow p + p + \bar{n}$ proceeds via the strong interaction. By applying the conservation laws, deduce the (a) charge quantum number, (b) baryon number, and (c) strangeness of the antineutron.

1246 CHAPTER 44 QUARKS, LEPTONS, AND THE BIG BANG

••20 **GO** Consider the decay $\Lambda^0 \rightarrow p + \pi^-$ with the Λ^0 at rest. (a) Calculate the disintegration energy. What is the kinetic energy of (b) the proton and (c) the pion? (*Hint:* See Problem 6.)

••21 Use the conservation laws and Tables 44-3 and 44-4 to identify particle x in each of the following reactions, which proceed by means of the strong interaction: (a) $p + p \rightarrow p + \Lambda^0 + x$; (b) $p + \bar{p} \rightarrow n + x$; (c) $\pi^- + p \rightarrow \Xi^0 + K^0 + x$.

••22 A Σ^- particle moving with 220 MeV of kinetic energy decays according to $\Sigma^- \rightarrow \pi^- + n$. Calculate the total kinetic energy of the decay products.

sec. 44-8 The Quark Model

••23 **SSM WWW** Using the up, down, and strange quarks only, construct, if possible, a baryon (a) with $q = +1$ and strangeness $S = -2$ and (b) with $q = +2$ and strangeness $S = 0$.

••24 There are 10 baryons with spin $\frac{3}{2}$. Their symbols and quantum numbers for charge q and strangeness S are as follows:

q	S	q	S		
Δ^-	-1	0	Σ^{*0}	0	-1
Δ^0	0	0	Σ^{*+}	+1	-1
Δ^+	+1	0	Ξ^{*-}	-1	-2
Δ^{++}	+2	0	Ξ^{*0}	0	-2
Σ^{*-}	-1	-1	Ω^-	-1	-3

Make a charge-strangeness plot for these baryons, using the sloping coordinate system of Fig. 44-3. Compare your plot with this figure.

••25 The quark makeups of the proton and neutron are uud and udd, respectively. What are the quark makeups of (a) the antiproton and (b) the antineutron?

••26 From Tables 44-3 and 44-5, determine the identity of the baryon formed from quarks (a) ddu, (b) uus, and (c) ssd. Check your answers against the baryon octet shown in Fig. 44-3a.

••27 What is the quark makeup of \bar{K}^0 ?

••28 What quark combination is needed to form (a) Λ^0 and (b) Ξ^0 ?

••29 Which hadron in Tables 44-3 and 44-4 corresponds to the quark bundles (a) ssu and (b) dds?

••30 The spin- $\frac{3}{2}$ Σ^{*0} baryon (see table in Problem 24) has a rest energy of 1385 MeV (with an intrinsic uncertainty ignored here); the spin- $\frac{1}{2}$ Σ^0 baryon has a rest energy of 1192.5 MeV. If each of these particles has a kinetic energy of 1000 MeV, (a) which is moving faster and (b) by how much?

sec. 44-11 The Universe Is Expanding

••31 In the laboratory, one of the lines of sodium is emitted at a wavelength of 590.0 nm. In the light from a particular galaxy, however, this line is seen at a wavelength of 602.0 nm. Calculate the distance to the galaxy, assuming that Hubble's law holds and that the Doppler shift of Eq. 37-36 applies.

••32 Because of the cosmological expansion, a particular emission from a distant galaxy has a wavelength that is 2.00 times the wavelength that emission would have in a laboratory. Assuming that Hubble's law holds and that we can apply Doppler-shift calculations, what was the distance (ly) to that galaxy when the light was emitted?

••33 What is the observed wavelength of the 656.3 nm (first Balmer) line of hydrogen emitted by a galaxy at a distance of

2.40×10^8 ly? Assume that the Doppler shift of Eq. 37-36 and Hubble's law apply.

••34 An object is 1.5×10^4 ly from us and does not have any motion relative to us except for the motion due to the expansion of the universe. If the space between us and it expands according to Hubble's law, with $H = 21.8$ mm/s · ly, (a) how much extra distance (meters) will be between us and the object by this time next year and (b) what is the speed of the object away from us?

••35 If Hubble's law can be extrapolated to very large distances, at what distance would the apparent recessional speed become equal to the speed of light?

••36 Will the universe continue to expand forever? To attack this question, assume that the theory of dark energy is in error and that the recessional speed v of a galaxy a distance r from us is determined only by the gravitational interaction of the matter that lies inside a sphere of radius r centered on us. If the total mass inside this sphere is M , the escape speed v_e from the sphere is $v_e = \sqrt{2GM/r}$ (Eq. 13-28). (a) Show that to prevent unlimited expansion, the average density ρ inside the sphere must be at least equal to

$$\rho = \frac{3H^2}{8\pi G}.$$

(b) Evaluate this "critical density" numerically; express your answer in terms of hydrogen atoms per cubic meter. Measurements of the actual density are difficult and are complicated by the presence of dark matter.

••37 Because the apparent recessional speeds of galaxies and quasars at great distances are close to the speed of light, the relativistic Doppler shift formula (Eq. 37-31) must be used. The shift is reported as fractional red shift $z = \Delta\lambda/\lambda_0$. (a) Show that, in terms of z , the recessional speed parameter $\beta = v/c$ is given by

$$\beta = \frac{z^2 + 2z}{z^2 + 2z + 2}.$$

(b) A quasar detected in 1987 has $z = 4.43$. Calculate its speed parameter. (c) Find the distance to the quasar, assuming that Hubble's law is valid to these distances.

••38 **GO** An electron jumps from $n = 3$ to $n = 2$ in a hydrogen atom in a distant galaxy, emitting light. If we detect that light at a wavelength of 3.00 mm, by what multiplication factor has the wavelength, and thus the universe, expanded since the light was emitted?

sec. 44-12 The Cosmic Background Radiation

••39 Due to the presence everywhere of the cosmic background radiation, the minimum possible temperature of a gas in interstellar or intergalactic space is not 0 K but 2.7 K. This implies that a significant fraction of the molecules in space that can be in a low-level excited state may, in fact, be so. Subsequent de-excitation would lead to the emission of radiation that could be detected. Consider a (hypothetical) molecule with just one possible excited state. (a) What would the excitation energy have to be for 25% of the molecules to be in the excited state? (*Hint:* See Eq. 40-29.) (b) What would be the wavelength of the photon emitted in a transition back to the ground state?

sec. 44-13 Dark Matter

••40 What would the mass of the Sun have to be if Pluto (the outermost planet most of the time) were to have the same orbital speed that Mercury (the innermost planet) has now? Use data

PROBLEMS

1247

from Appendix C, express your answer in terms of the Sun's current mass M_S , and assume circular orbits.

••41 SSM Suppose that the radius of the Sun were increased to 5.90×10^{12} m (the average radius of the orbit of Pluto), that the density of this expanded Sun were uniform, and that the planets revolved within this tenuous object. (a) Calculate Earth's orbital speed in this new configuration. (b) What is the ratio of the orbital speed calculated in (a) to Earth's present orbital speed of 29.8 km/s? Assume that the radius of Earth's orbit remains unchanged. (c) What would be Earth's new period of revolution? (The Sun's mass remains unchanged.)

••42 Suppose that the matter (stars, gas, dust) of a particular galaxy, of total mass M , is distributed uniformly throughout a sphere of radius R . A star of mass m is revolving about the center of the galaxy in a circular orbit of radius $r < R$. (a) Show that the orbital speed v of the star is given by

$$v = r \sqrt{GM/R^3},$$

and therefore that the star's period T of revolution is

$$T = 2\pi \sqrt{R^3/GM},$$

independent of r . Ignore any resistive forces. (b) Next suppose that the galaxy's mass is concentrated near the galactic center, within a sphere of radius less than r . What expression then gives the star's orbital period?

sec. 44-14 The Big Bang

•43 The wavelength at which a thermal radiator at temperature T radiates electromagnetic waves most intensely is given by Wien's law: $\lambda_{\max} = (2898 \mu\text{m} \cdot \text{K})/T$. (a) Show that the energy E of a photon corresponding to that wavelength can be computed from

$$E = (4.28 \times 10^{-10} \text{ MeV/K})T.$$

(b) At what minimum temperature can this photon create an electron–positron pair (as discussed in Section 21-6)?

•44 Use Wien's law (see Problem 43) to answer the following questions: (a) The cosmic background radiation peaks in intensity at a wavelength of 1.1 mm. To what temperature does this correspond? (b) About 379 000 y after the big bang, the universe became transparent to electromagnetic radiation. Its temperature then was 2970 K. What was the wavelength at which the background radiation was then most intense?

Additional Problems

45 SSM There is no known meson with charge quantum number $q = +1$ and strangeness $S = -1$ or with $q = -1$ and $S = +1$. Explain why in terms of the quark model.

46 Figure 44-11 is a hypothetical plot of the recessional speeds v of galaxies against their distance r from us; the best-fit straight line through the data points is shown. From this plot determine the age of the universe, assuming that Hubble's law holds and that Hubble's constant has always had the same value.

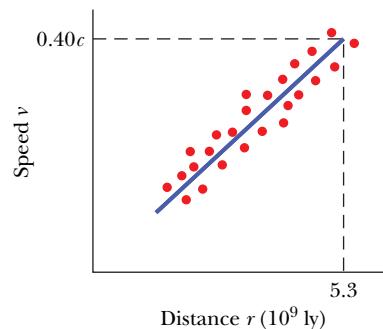


Fig. 44-11 Problem 46.

47 SSM How much energy would be released if Earth were annihilated by collision with an anti-Earth?

48 *A particle game.* Figure 44-12 is a sketch of the tracks made by particles in a *fictional* cloud chamber experiment (with a uniform magnetic field directed perpendicular to the page), and Table 44-6 gives *fictional* quantum numbers associated with the particles making the tracks. Particle *A* entered the chamber at the lower left, leaving track 1 and decaying into three particles. Then the particle creating track 6 decayed into three other particles, and the particle creating track 4 decayed into two other particles, one of which was electrically uncharged—the path of that uncharged particle is represented by the dashed straight line because, being electrically neutral, it would not actually leave a track in a cloud chamber. The particle that created track 8 is known to have a seriousness quantum number of zero.

By conserving the fictional quantum numbers at each decay point and by noting the directions of curvature of the tracks, identify which particle goes with track (a) 1, (b) 2, (c) 3, (d) 4, (e) 5, (f) 6, (g) 7, (h) 8, and (i) 9. One of the listed particles is not formed; the others appear only once each.

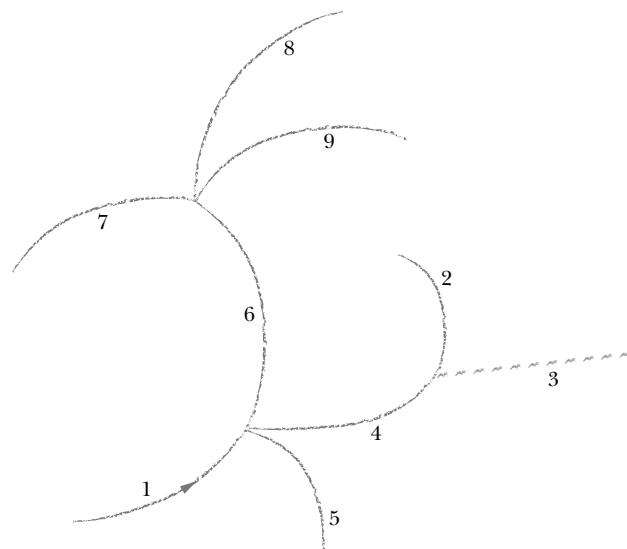


Fig. 44-12 Problem 48.

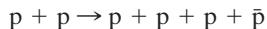
Table 44-6

Problem 44-48

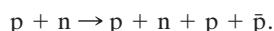
Particle	Charge	Whimsy	Seriousness	Cuteness
<i>A</i>	1	1	-2	-2
<i>B</i>	0	4	3	0
<i>C</i>	1	2	-3	-1
<i>D</i>	-1	-1	0	1
<i>E</i>	-1	0	-4	-2
<i>F</i>	1	0	0	0
<i>G</i>	-1	-1	1	-1
<i>H</i>	3	3	1	0
<i>I</i>	0	6	4	6
<i>J</i>	1	-6	-4	-6

1248 CHAPTER 44 QUARKS, LEPTONS, AND THE BIG BANG

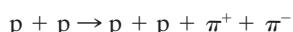
49 Figure 44-13 shows part of the experimental arrangement in which antiprotons were discovered in the 1950s. A beam of 6.2 GeV protons emerged from a particle accelerator and collided with nuclei in a copper target. According to theoretical predictions at the time, collisions between protons in the beam and the protons and neutrons in those nuclei should produce antiprotons via the reactions



and



However, even if these reactions did occur, they would be rare compared to the reactions



and

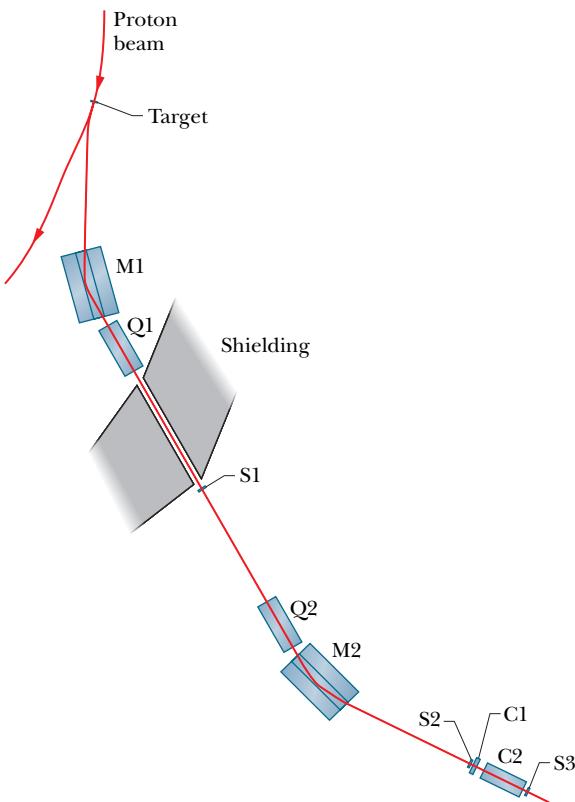
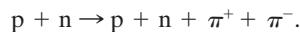


Fig. 44-13 Problem 49.

Thus, most of the particles produced by the collisions between the 6.2 GeV protons and the copper target were pions.

To prove that antiprotons exist and were produced by some limited number of the collisions, particles leaving the target were sent into a series of magnetic fields and detectors as shown in Fig. 44-13. The first magnetic field (M1) curved the path of any charged particle passing through it; moreover, the field was arranged so that the only particles that emerged from it to reach the second magnetic field (Q1) had to be negatively charged (either a \bar{p} or a π^-) and have a momentum of 1.19 GeV/c. Field Q1 was a special type of magnetic field (a *quadrupole field*) that focused the particles reaching it into a beam, allowing them to pass through a hole in thick shielding to a *scintillation counter* S1. The passage of a charged particle through the counter triggered a signal, with each signal indicating the passage of either a 1.19 GeV/c π^- or (presumably) a 1.19 GeV/c \bar{p} .

After being refocused by magnetic field Q2, the particles were directed by magnetic field M2 through a second scintillation counter S2 and then through two *Cerenkov counters* C1 and C2. These latter detectors can be manufactured so that they send a signal only when the particle passing through them is moving with a speed that falls within a certain range. In the experiment, a particle with a speed greater than $0.79c$ would trigger C1 and a particle with a speed between $0.75c$ and $0.78c$ would trigger C2.

There were then two ways to distinguish the predicted rare antiprotons from the abundant negative pions. Both ways involved the fact that the speed of a 1.19 GeV/c \bar{p} differs from that of a 1.19 GeV/c π^- : (1) According to calculations, a \bar{p} would trigger one of the Cerenkov counters and a π^- would trigger the other. (2) The time interval Δt between signals from S1 and S2, which were separated by 12 m, would have one value for a \bar{p} and another value for a π^- . Thus, if the correct Cerenkov counter were triggered and the time interval Δt had the correct value, the experiment would prove the existence of antiprotons.

What is the speed of (a) an antiproton with a momentum of 1.19 GeV/c and (b) a negative pion with that same momentum? (The speed of an antiproton through the Cerenkov detectors would actually be slightly less than calculated here because the antiproton would lose a little energy within the detectors.) Which Cerenkov detector was triggered by (c) an antiproton and (d) a negative pion? What time interval Δt indicated the passage of (e) an antiproton and (f) a negative pion? [Problem adapted from O. Chamberlain, E. Segrè, C. Wiegand, and T. Ypsilantis, "Observation of Antiprotons," *Physical Review*, Vol. 100, pp. 947–950 (1955).]

50 Verify that the hypothetical proton decay scheme in Eq. 44-14 does not violate the conservation law of (a) charge, (b) energy, and (c) linear momentum. (d) How about angular momentum?

The International System of Units (SI)*

TABLE 1

The SI Base Units

Quantity	Name	Symbol	Definition
length	meter	m	“... the length of the path traveled by light in vacuum in 1/299,792,458 of a second.” (1983)
mass	kilogram	kg	“... this prototype [a certain platinum–iridium cylinder] shall henceforth be considered to be the unit of mass.” (1889)
time	second	s	“... the duration of 9,192,631,770 periods of the radiation corresponding to the transition between the two hyperfine levels of the ground state of the cesium-133 atom.” (1967)
electric current	ampere	A	“... that constant current which, if maintained in two straight parallel conductors of infinite length, of negligible circular cross section, and placed 1 meter apart in vacuum, would produce between these conductors a force equal to 2×10^{-7} newton per meter of length.” (1946)
thermodynamic temperature	kelvin	K	“... the fraction 1/273.16 of the thermodynamic temperature of the triple point of water.” (1967)
amount of substance	mole	mol	“... the amount of substance of a system which contains as many elementary entities as there are atoms in 0.012 kilogram of carbon-12.” (1971)
luminous intensity	candela	cd	“... the luminous intensity, in a given direction, of a source that emits monochromatic radiation of frequency 540×10^{12} hertz and that has a radiant intensity in that direction of 1/683 watt per steradian.” (1979)

*Adapted from “The International System of Units (SI),” National Bureau of Standards Special Publication 330, 1972 edition. The definitions above were adopted by the General Conference of Weights and Measures, an international body, on the dates shown. In this book we do not use the candela.

A-2 APPENDIX A THE INTERNATIONAL SYSTEM OF UNITS (SI)

TABLE 2

Some SI Derived Units

Quantity	Name of Unit	Symbol
area	square meter	m^2
volume	cubic meter	m^3
frequency	hertz	Hz
mass density (density)	kilogram per cubic meter	kg/m^3
speed, velocity	meter per second	m/s
angular velocity	radian per second	rad/s
acceleration	meter per second per second	m/s^2
angular acceleration	radian per second per second	rad/s^2
force	newton	N
pressure	pascal	Pa
work, energy, quantity of heat	joule	J
power	watt	W
quantity of electric charge	coulomb	C
potential difference, electromotive force	volt	V
electric field strength	volt per meter (or newton per coulomb)	V/m
electric resistance	ohm	Ω
capacitance	farad	F
magnetic flux	weber	Wb
inductance	henry	H
magnetic flux density	tesla	T
magnetic field strength	ampere per meter	A/m
entropy	joule per kelvin	J/K
specific heat	joule per kilogram kelvin	$\text{J}/(\text{kg} \cdot \text{K})$
thermal conductivity	watt per meter kelvin	$\text{W}/(\text{m} \cdot \text{K})$
radiant intensity	watt per steradian	W/sr

TABLE 3

The SI Supplementary Units

Quantity	Name of Unit	Symbol
plane angle	radian	rad
solid angle	steradian	sr

Some Fundamental Constants of Physics*

Constant	Symbol	Computational Value	Best (1998) Value	
			Value ^a	Uncertainty ^b
Speed of light in a vacuum	c	3.00×10^8 m/s	2.997 924 58	exact
Elementary charge	e	1.60×10^{-19} C	1.602 176 487	0.025
Gravitational constant	G	6.67×10^{-11} m ³ /s ² · kg	6.674 28	100
Universal gas constant	R	8.31 J/mol · K	8.314 472	1.7
Avogadro constant	N_A	6.02×10^{23} mol ⁻¹	6.022 141 79	0.050
Boltzmann constant	k	1.38×10^{-23} J/K	1.380 650 4	1.7
Stefan–Boltzmann constant	σ	5.67×10^{-8} W/m ² · K ⁴	5.670 400	7.0
Molar volume of ideal gas at STP ^d	V_m	2.27×10^{-2} m ³ /mol	2.271 098 1	1.7
Permittivity constant	ϵ_0	8.85×10^{-12} F/m	8.854 187 817 62	exact
Permeability constant	μ_0	1.26×10^{-6} H/m	1.256 637 061 43	exact
Planck constant	h	6.63×10^{-34} J · s	6.626 068 96	0.050
Electron mass ^c	m_e	9.11×10^{-31} kg 5.49×10^{-4} u	9.109 382 15 5.485 799 094 3	0.050 4.2×10^{-4}
Proton mass ^c	m_p	1.67×10^{-27} kg 1.0073 u	1.672 621 637 1.007 276 466 77	0.050 1.0×10^{-4}
Ratio of proton mass to electron mass	m_p/m_e	1840	1836.152 672 47	4.3×10^{-4}
Electron charge-to-mass ratio	e/m_e	1.76×10^{11} C/kg	1.758 820 150	0.025
Neutron mass ^c	m_n	1.68×10^{-27} kg 1.0087 u	1.674 927 211 1.008 664 915 97	0.050 4.3×10^{-4}
Hydrogen atom mass ^c	m_{1_H}	1.0078 u	1.007 825 031 6	0.0005
Deuterium atom mass ^c	m_{2_H}	2.0136 u	2.013 553 212 724	3.9×10^{-5}
Helium atom mass ^c	$m_{4_{He}}$	4.0026 u	4.002 603 2	0.067
Muon mass	m_μ	1.88×10^{-28} kg	1.883 531 30	0.056
Electron magnetic moment	μ_e	9.28×10^{-24} J/T	9.284 763 77	0.025
Proton magnetic moment	μ_p	1.41×10^{-26} J/T	1.410 606 662	0.026
Bohr magneton	μ_B	9.27×10^{-24} J/T	9.274 009 15	0.025
Nuclear magneton	μ_N	5.05×10^{-27} J/T	5.050 783 24	0.025
Bohr radius	a	5.29×10^{-11} m	5.291 772 085 9	6.8×10^{-4}
Rydberg constant	R	1.10×10^7 m ⁻¹	1.097 373 156 852 7	6.6×10^{-6}
Electron Compton wavelength	λ_C	2.43×10^{-12} m	2.426 310 217 5	0.0014

^aValues given in this column should be given the same unit and power of 10 as the computational value.

^bParts per million.

^cMasses given in u are in unified atomic mass units, where 1 u = $1.660\ 538\ 782 \times 10^{-27}$ kg.

^dSTP means standard temperature and pressure: 0°C and 1.0 atm (0.1 MPa).

*The values in this table were selected from the 1998 CODATA recommended values (www.physics.nist.gov).

Some Astronomical Data

Some Distances from Earth

To the Moon*	3.82×10^8 m	To the center of our galaxy	2.2×10^{20} m
To the Sun*	1.50×10^{11} m	To the Andromeda Galaxy	2.1×10^{22} m
To the nearest star (Proxima Centauri)	4.04×10^{16} m	To the edge of the observable universe	$\sim 10^{26}$ m

*Mean distance.

The Sun, Earth, and the Moon

Property	Unit	Sun	Earth	Moon
Mass	kg	1.99×10^{30}	5.98×10^{24}	7.36×10^{22}
Mean radius	m	6.96×10^8	6.37×10^6	1.74×10^6
Mean density	kg/m ³	1410	5520	3340
Free-fall acceleration at the surface	m/s ²	274	9.81	1.67
Escape velocity	km/s	618	11.2	2.38
Period of rotation ^a	—	37 d at poles ^b 26 d at equator ^b	23 h 56 min	27.3 d
Radiation power ^c	W	3.90×10^{26}		

^aMeasured with respect to the distant stars.

^bThe Sun, a ball of gas, does not rotate as a rigid body.

^cJust outside Earth's atmosphere solar energy is received, assuming normal incidence, at the rate of 1340 W/m².

Some Properties of the Planets

	Mercury	Venus	Earth	Mars	Jupiter	Saturn	Uranus	Neptune	Pluto
Mean distance from Sun, 10^6 km	57.9	108	150	228	778	1430	2870	4500	5900
Period of revolution, y	0.241	0.615	1.00	1.88	11.9	29.5	84.0	165	248
Period of rotation, ^a d	58.7	-243 ^b	0.997	1.03	0.409	0.426	-0.451 ^b	0.658	6.39
Orbital speed, km/s	47.9	35.0	29.8	24.1	13.1	9.64	6.81	5.43	4.74
Inclination of axis to orbit	<28°	≈3°	23.4°	25.0°	3.08°	26.7°	97.9°	29.6°	57.5°
Inclination of orbit to Earth's orbit	7.00°	3.39°		1.85°	1.30°	2.49°	0.77°	1.77°	17.2°
Eccentricity of orbit	0.206	0.0068	0.0167	0.0934	0.0485	0.0556	0.0472	0.0086	0.250
Equatorial diameter, km	4880	12 100	12 800	6790	143 000	120 000	51 800	49 500	2300
Mass (Earth = 1)	0.0558	0.815	1.000	0.107	318	95.1	14.5	17.2	0.002
Density (water = 1)	5.60	5.20	5.52	3.95	1.31	0.704	1.21	1.67	2.03
Surface value of g , ^c m/s ²	3.78	8.60	9.78	3.72	22.9	9.05	7.77	11.0	0.5
Escape velocity, ^c km/s	4.3	10.3	11.2	5.0	59.5	35.6	21.2	23.6	1.3
Known satellites	0	0	1	2	63 + ring	60 + rings	27 + rings	13 + rings	3

^aMeasured with respect to the distant stars.

^bVenus and Uranus rotate opposite their orbital motion.

^cGravitational acceleration measured at the planet's equator.

Conversion Factors

Conversion factors may be read directly from these tables. For example, 1 degree = 2.778×10^{-3} revolutions, so $16.7^\circ = 16.7 \times 2.778 \times 10^{-3}$ rev. The SI units are fully capitalized. Adapted in part from G. Shortley and D. Williams, *Elements of Physics*, 1971, Prentice-Hall, Englewood Cliffs, NJ.

Plane Angle

	$^\circ$	'	"	RADIAN	rev
1 degree = 1	60		3600	1.745×10^{-2}	2.778×10^{-3}
1 minute = 1.667×10^{-2}	1		60	2.909×10^{-4}	4.630×10^{-5}
1 second = 2.778×10^{-4}		1.667×10^{-2}	1	4.848×10^{-6}	7.716×10^{-7}
1 RADIAN = 57.30	3438		2.063×10^5	1	0.1592
1 revolution = 360		2.16×10^4	1.296×10^6	6.283	1

Solid Angle

$$1 \text{ sphere} = 4\pi \text{ steradians} = 12.57 \text{ steradians}$$

Length

	cm	METER	km	in.	ft	mi
1 centimeter = 1	10^{-2}	10^{-5}	0.3937	3.281×10^{-2}	6.214×10^{-6}	
1 METER = 100	1	10^{-3}	39.37	3.281	6.214×10^{-4}	
1 kilometer = 10^5	1000	1	3.937×10^4	3281	0.6214	
1 inch = 2.540		2.540×10^{-2}	2.540×10^{-5}	1	8.333×10^{-2}	1.578×10^{-5}
1 foot = 30.48	0.3048		3.048×10^{-4}	12	1	1.894×10^{-4}
1 mile = 1.609×10^5	1609	1.609	6.336×10^4	5280	1	
1 angström = 10^{-10} m		1 fermi = 10^{-15} m		1 fathom = 6 ft		1 rod = 16.5 ft
1 nautical mile = 1852 m = 1.151 miles = 6076 ft		1 light-year = 9.461×10^{12} km 1 parsec = 3.084×10^{13} km		1 Bohr radius = 5.292×10^{-11} m 1 yard = 3 ft		1 mil = 10^{-3} in. 1 nm = 10^{-9} m

Area

	METER ²	cm ²	ft ²	in. ²
1 SQUARE METER = 1	10^4	10.76	1550	
1 square centimeter = 10^{-4}	1	1.076×10^{-3}	0.1550	
1 square foot = 9.290×10^{-2}	929.0	1	144	
1 square inch = 6.452×10^{-4}	6.452	6.944×10^{-3}	1	

$$1 \text{ square mile} = 2.788 \times 10^7 \text{ ft}^2 = 640 \text{ acres}$$

$$1 \text{ barn} = 10^{-28} \text{ m}^2$$

$$1 \text{ acre} = 43560 \text{ ft}^2$$

$$1 \text{ hectare} = 10^4 \text{ m}^2 = 2.471 \text{ acres}$$

A-6 APPENDIX D CONVERSION FACTORS

Volume

METER ³	cm ³	L	ft ³	in. ³
1 CUBIC METER = 1	10 ⁶	1000	35.31	6.102 × 10 ⁴
1 cubic centimeter = 10 ⁻⁶	1	1.000 × 10 ⁻³	3.531 × 10 ⁻⁵	6.102 × 10 ⁻²
1 liter = 1.000 × 10 ⁻³	1000	1	3.531 × 10 ⁻²	61.02
1 cubic foot = 2.832 × 10 ⁻²	2.832 × 10 ⁴	28.32	1	1728
1 cubic inch = 1.639 × 10 ⁻⁵	16.39	1.639 × 10 ⁻²	5.787 × 10 ⁻⁴	1

1 U.S. fluid gallon = 4 U.S. fluid quarts = 8 U.S. pints = 128 U.S. fluid ounces = 231 in.³

1 British imperial gallon = 277.4 in.³ = 1.201 U.S. fluid gallons

Mass

Quantities in the colored areas are not mass units but are often used as such. For example, when we write 1 kg “=” 2.205 lb, this means that a kilogram is a *mass* that *weighs* 2.205 pounds at a location where *g* has the standard value of 9.80665 m/s².

g	KILOGRAM	slug	u	oz	lb	ton
1 gram = 1	0.001	6.852 × 10 ⁻⁵	6.022 × 10 ²³	3.527 × 10 ⁻²	2.205 × 10 ⁻³	1.102 × 10 ⁻⁶
1 KILOGRAM = 1000	1	6.852 × 10 ⁻²	6.022 × 10 ²⁶	35.27	2.205	1.102 × 10 ⁻³
1 slug = 1.459 × 10 ⁴	14.59	1	8.786 × 10 ²⁷	514.8	32.17	1.609 × 10 ⁻²
1 atomic mass unit = 1.661 × 10 ⁻²⁴	1.661 × 10 ⁻²⁷	1.138 × 10 ⁻²⁸	1	5.857 × 10 ⁻²⁶	3.662 × 10 ⁻²⁷	1.830 × 10 ⁻³⁰
1 ounce = 28.35	2.835 × 10 ⁻²	1.943 × 10 ⁻³	1.718 × 10 ²⁵	1	6.250 × 10 ⁻²	3.125 × 10 ⁻⁵
1 pound = 453.6	0.4536	3.108 × 10 ⁻²	2.732 × 10 ²⁶	16	1	0.0005
1 ton = 9.072 × 10 ⁵	907.2	62.16	5.463 × 10 ²⁹	3.2 × 10 ⁴	2000	1

1 metric ton = 1000 kg

Density

Quantities in the colored areas are weight densities and, as such, are dimensionally different from mass densities. See the note for the mass table.

slug/ft ³	KILOGRAM/ METER ³	g/cm ³	lb/ft ³	lb/in. ³
1 slug per foot ³ = 1	515.4	0.5154	32.17	1.862 × 10 ⁻²
1 KILOGRAM per METER ³ = 1.940 × 10 ⁻³	1	0.001	6.243 × 10 ⁻²	3.613 × 10 ⁻⁵
1 gram per centimeter ³ = 1.940	1000	1	62.43	3.613 × 10 ⁻²
1 pound per foot ³ = 3.108 × 10 ⁻²	16.02	16.02 × 10 ⁻²	1	5.787 × 10 ⁻⁴
1 pound per inch ³ = 53.71	2.768 × 10 ⁴	27.68	1728	1

Time

y	d	h	min	SECOND
1 year = 1	365.25	8.766 × 10 ³	5.259 × 10 ⁵	3.156 × 10 ⁷
1 day = 2.738 × 10 ⁻³	1	24	1440	8.640 × 10 ⁴
1 hour = 1.141 × 10 ⁻⁴	4.167 × 10 ⁻²	1	60	3600
1 minute = 1.901 × 10 ⁻⁶	6.944 × 10 ⁻⁴	1.667 × 10 ⁻²	1	60
1 SECOND = 3.169 × 10 ⁻⁸	1.157 × 10 ⁻⁵	2.778 × 10 ⁻⁴	1.667 × 10 ⁻²	1

Speed

	ft/s	km/h	METER/SECOND	mi/h	cm/s
1 foot per second = 1		1.097	0.3048	0.6818	30.48
1 kilometer per hour = 0.9113		1	0.2778	0.6214	27.78
1 METER per SECOND = 3.281		3.6	1	2.237	100
1 mile per hour = 1.467		1.609	0.4470	1	44.70
1 centimeter per second = 3.281×10^{-2}		3.6×10^{-2}	0.01	2.237×10^{-2}	1

1 knot = 1 nautical mi/h = 1.688 ft/s 1 mi/min = 88.00 ft/s = 60.00 mi/h

Force

Force units in the colored areas are now little used. To clarify: 1 gram-force (= 1 gf) is the force of gravity that would act on an object whose mass is 1 gram at a location where g has the standard value of 9.80665 m/s^2 .

	dyne	NEWTON	lb	pdl	gf	kgf
1 dyne = 1	10^{-5}		2.248×10^{-6}	7.233×10^{-5}	1.020×10^{-3}	1.020×10^{-6}
1 NEWTON = 10^5	1		0.2248	7.233	102.0	0.1020
1 pound = 4.448×10^5	4.448		1	32.17	453.6	0.4536
1 poundal = 1.383×10^4	0.1383		3.108×10^{-2}	1	14.10	1.410×10^2
1 gram-force = 980.7	9.807×10^{-3}		2.205×10^{-3}	7.093×10^{-2}	1	0.001
1 kilogram-force = 9.807×10^5	9.807		2.205	70.93	1000	1

1 ton = 2000 lb

Pressure

	atm	dyne/cm ²	inch of water	cm Hg	PASCAL	lb/in. ²	lb/ft ²
1 atmosphere = 1		1.013×10^6	406.8	76	1.013×10^5	14.70	2116
1 dyne per centimeter ² = 9.869×10^{-7}	1		4.015×10^{-4}	7.501×10^{-5}	0.1	1.405×10^{-5}	2.089×10^{-3}
1 inch of water ^a at 4°C = 2.458×10^{-3}	2491		1	0.1868	249.1	3.613×10^{-2}	5.202
1 centimeter of mercury ^a							
at 0°C = 1.316×10^{-2}		1.333×10^4	5.353	1	1333	0.1934	27.85
1 PASCAL = 9.869×10^{-6}	10		4.015×10^{-3}	7.501×10^{-4}	1	1.450×10^{-4}	2.089×10^{-2}
1 pound per inch ² = 6.805×10^{-2}		6.895×10^4	27.68	5.171	6.895×10^3	1	144
1 pound per foot ² = 4.725×10^{-4}	478.8		0.1922	3.591×10^{-2}	47.88	6.944×10^{-3}	1

^aWhere the acceleration of gravity has the standard value of 9.80665 m/s^2 .1 bar = 10^6 dyne/cm² = 0.1 MPa1 millibar = 10^3 dyne/cm² = 10^2 Pa

1 torr = 1 mm Hg

A-8 APPENDIX D CONVERSION FACTORS

Energy, Work, Heat

Quantities in the colored areas are not energy units but are included for convenience. They arise from the relativistic mass-energy equivalence formula $E = mc^2$ and represent the energy released if a kilogram or unified atomic mass unit (u) is completely converted to energy (bottom two rows) or the mass that would be completely converted to one unit of energy (rightmost two columns).

	Btu	erg	ft · lb	hp · h	JOULE	cal	kW · h	eV	MeV	kg	u	
1 British thermal unit = 1	1.055 $\times 10^{10}$	777.9	3.929 $\times 10^{-4}$	1055	252.0	2.930 $\times 10^{-4}$	6.585 $\times 10^{21}$	6.585 $\times 10^{15}$	1.174 $\times 10^{-14}$	7.070 $\times 10^{12}$		
1 erg = $\times 10^{-11}$	9.481	1	7.376 $\times 10^{-8}$	3.725 $\times 10^{-14}$	10 ⁻⁷	2.389 $\times 10^{-8}$	2.778 $\times 10^{-14}$	6.242 $\times 10^{11}$	6.242 $\times 10^5$	1.113 $\times 10^{-24}$	670.2	
1 foot-pound = $\times 10^{-3}$	1.285	1.356 $\times 10^7$	1	$\times 10^{-7}$	5.051	1.356	0.3238 $\times 10^{-7}$	3.766 $\times 10^{-7}$	8.464 $\times 10^{18}$	8.464 $\times 10^{12}$	1.509 $\times 10^{-17}$	9.037 $\times 10^9$
1 horsepower-hour = 2545	2.685 $\times 10^{13}$	1.980 $\times 10^6$	1		2.685 $\times 10^6$	6.413 $\times 10^5$	0.7457 $\times 10^{25}$	1.676 $\times 10^{19}$	1.676 $\times 10^{19}$	2.988 $\times 10^{-11}$	1.799 $\times 10^{16}$	
1 JOULE = $\times 10^{-4}$	9.481	10 ⁷	0.7376	$\times 10^{-7}$	1	0.2389	2.778 $\times 10^{-7}$	6.242 $\times 10^{18}$	6.242 $\times 10^{12}$	1.113 $\times 10^{-17}$	6.702 $\times 10^9$	
1 calorie = $\times 10^{-3}$	3.968	4.1868 $\times 10^7$	3.088 $\times 10^{-6}$	1.560	4.1868	1	1.163 $\times 10^{-6}$	2.613 $\times 10^{19}$	2.613 $\times 10^{13}$	4.660 $\times 10^{-17}$	2.806 $\times 10^{10}$	
1 kilowatt-hour = 3413	3.600 $\times 10^{13}$	2.655 $\times 10^6$	1.341 $\times 10^6$		3.600 $\times 10^5$	8.600 $\times 10^5$	1	2.247 $\times 10^{25}$	2.247 $\times 10^{19}$	4.007 $\times 10^{-11}$	2.413 $\times 10^{16}$	
1 electron-volt = $\times 10^{-22}$	1.519	1.602 $\times 10^{-12}$	1.182 $\times 10^{-19}$	5.967 $\times 10^{-26}$	1.602 $\times 10^{-19}$	3.827 $\times 10^{-20}$	4.450 $\times 10^{-26}$	1	10 ⁻⁶	1.783 $\times 10^{-36}$	1.074 $\times 10^{-9}$	
1 million electron-volts = $\times 10^{-16}$	1.519	1.602 $\times 10^{-6}$	1.182 $\times 10^{-13}$	5.967 $\times 10^{-20}$	1.602 $\times 10^{-13}$	3.827 $\times 10^{-14}$	4.450 $\times 10^{-20}$	10 ⁻⁶	1	1.783 $\times 10^{-30}$	1.074 $\times 10^{-3}$	
1 kilogram = $\times 10^{13}$	8.521	8.987 $\times 10^{23}$	6.629 $\times 10^{16}$	3.348 $\times 10^{10}$	8.987 $\times 10^{16}$	2.146 $\times 10^{16}$	2.497 $\times 10^{10}$	5.610 $\times 10^{35}$	5.610 $\times 10^{29}$	1	6.022 $\times 10^{26}$	
1 unified atomic mass unit = $\times 10^{-13}$	1.415	1.492 $\times 10^{-3}$	1.101 $\times 10^{-10}$	5.559 $\times 10^{-17}$	1.492 $\times 10^{-10}$	3.564 $\times 10^{-11}$	4.146 $\times 10^{-17}$	9.320 $\times 10^8$	932.0	1.661 $\times 10^{-27}$	1	

Power

	Btu/h	ft · lb/s	hp	cal/s	kW	WATT
1 British thermal unit per hour = 1	0.2161	3.929×10^{-4}	6.998×10^{-2}	2.930×10^{-4}	0.2930	
1 foot-pound per second = 4.628	1	1.818×10^{-3}	0.3239	1.356×10^{-3}	1.356	
1 horsepower = 2545	550	1	178.1	0.7457	745.7	
1 calorie per second = 14.29	3.088	5.615×10^{-3}	1	4.186×10^{-3}	4.186	
1 kilowatt = 3413	737.6	1.341	238.9	1	1000	
1 WATT = 3.413	0.7376	1.341×10^{-3}	0.2389	0.001	1	

Magnetic Field

gauss	TESLA	milligauss
1 gauss = 1	10^{-4}	1000
1 TESLA = 10^4	1	10^7
1 milligauss = 0.001	10^{-7}	1

Magnetic Flux

maxwell	WEBER
1 maxwell = 1	10^{-8}
1 WEBER = 10^8	1

1 tesla = 1 weber/meter²

Mathematical Formulas

Geometry

Circle of radius r : circumference = $2\pi r$; area = πr^2 .

Sphere of radius r : area = $4\pi r^2$; volume = $\frac{4}{3}\pi r^3$.

Right circular cylinder of radius r and height h :
area = $2\pi r^2 + 2\pi rh$; volume = $\pi r^2 h$.

Triangle of base a and altitude h : area = $\frac{1}{2}ah$.

Quadratic Formula

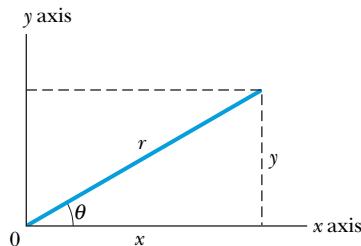
If $ax^2 + bx + c = 0$, then $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$.

Trigonometric Functions of Angle θ

$$\sin \theta = \frac{y}{r} \quad \cos \theta = \frac{x}{r}$$

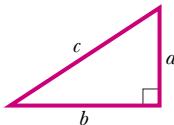
$$\tan \theta = \frac{y}{x} \quad \cot \theta = \frac{x}{y}$$

$$\sec \theta = \frac{r}{x} \quad \csc \theta = \frac{r}{y}$$



Pythagorean Theorem

In this right triangle,
 $a^2 + b^2 = c^2$



Triangles

Angles are A, B, C

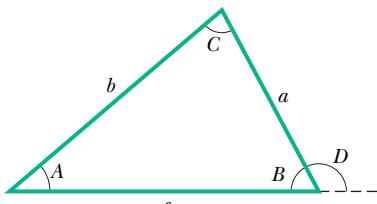
Opposite sides are a, b, c

Angles $A + B + C = 180^\circ$

$$\frac{\sin A}{a} = \frac{\sin B}{b} = \frac{\sin C}{c}$$

$$c^2 = a^2 + b^2 - 2ab \cos C$$

Exterior angle $D = A + C$



Mathematical Signs and Symbols

= equals

\approx equals approximately

\sim is the order of magnitude of

\neq is not equal to

\equiv is identical to, is defined as

$>$ is greater than (\gg is much greater than)

$<$ is less than (\ll is much less than)

\geq is greater than or equal to (or, is no less than)

\leq is less than or equal to (or, is no more than)

\pm plus or minus

\propto is proportional to

Σ the sum of

x_{avg} the average value of x

Trigonometric Identities

$$\sin(90^\circ - \theta) = \cos \theta$$

$$\cos(90^\circ - \theta) = \sin \theta$$

$$\sin \theta / \cos \theta = \tan \theta$$

$$\sin^2 \theta + \cos^2 \theta = 1$$

$$\sec^2 \theta - \tan^2 \theta = 1$$

$$\csc^2 \theta - \cot^2 \theta = 1$$

$$\sin 2\theta = 2 \sin \theta \cos \theta$$

$$\cos 2\theta = \cos^2 \theta - \sin^2 \theta = 2 \cos^2 \theta - 1 = 1 - 2 \sin^2 \theta$$

$$\sin(\alpha \pm \beta) = \sin \alpha \cos \beta \pm \cos \alpha \sin \beta$$

$$\cos(\alpha \pm \beta) = \cos \alpha \cos \beta \mp \sin \alpha \sin \beta$$

$$\tan(\alpha \pm \beta) = \frac{\tan \alpha \pm \tan \beta}{1 \mp \tan \alpha \tan \beta}$$

$$\sin \alpha \pm \sin \beta = 2 \sin \frac{1}{2}(\alpha \pm \beta) \cos \frac{1}{2}(\alpha \mp \beta)$$

$$\cos \alpha + \cos \beta = 2 \cos \frac{1}{2}(\alpha + \beta) \cos \frac{1}{2}(\alpha - \beta)$$

$$\cos \alpha - \cos \beta = -2 \sin \frac{1}{2}(\alpha + \beta) \sin \frac{1}{2}(\alpha - \beta)$$

Binomial Theorem

$$(1 + x)^n = 1 + \frac{nx}{1!} + \frac{n(n-1)x^2}{2!} + \dots \quad (x^2 < 1)$$

Exponential Expansion

$$e^x = 1 + x + \frac{x^2}{2!} + \frac{x^3}{3!} + \dots$$

Logarithmic Expansion

$$\ln(1 + x) = x - \frac{1}{2}x^2 + \frac{1}{3}x^3 - \dots \quad (|x| < 1)$$

Trigonometric Expansions (θ in radians)

$$\sin \theta = \theta - \frac{\theta^3}{3!} + \frac{\theta^5}{5!} - \dots$$

$$\cos \theta = 1 - \frac{\theta^2}{2!} + \frac{\theta^4}{4!} - \dots$$

$$\tan \theta = \theta + \frac{\theta^3}{3} + \frac{2\theta^5}{15} + \dots$$

Cramer's Rule

Two simultaneous equations in unknowns x and y ,

$$a_1x + b_1y = c_1 \quad \text{and} \quad a_2x + b_2y = c_2,$$

have the solutions

$$x = \frac{\begin{vmatrix} c_1 & b_1 \\ c_2 & b_2 \end{vmatrix}}{\begin{vmatrix} a_1 & b_1 \\ a_2 & b_2 \end{vmatrix}} = \frac{c_1b_2 - c_2b_1}{a_1b_2 - a_2b_1}$$

and

$$y = \frac{\begin{vmatrix} a_1 & c_1 \\ a_2 & c_2 \end{vmatrix}}{\begin{vmatrix} a_1 & b_1 \\ a_2 & b_2 \end{vmatrix}} = \frac{a_1c_2 - a_2c_1}{a_1b_2 - a_2b_1}.$$

Products of Vectors

Let \hat{i} , \hat{j} , and \hat{k} be unit vectors in the x , y , and z directions. Then

$$\hat{i} \cdot \hat{i} = \hat{j} \cdot \hat{j} = \hat{k} \cdot \hat{k} = 1, \quad \hat{i} \cdot \hat{j} = \hat{j} \cdot \hat{k} = \hat{k} \cdot \hat{i} = 0,$$

$$\hat{i} \times \hat{i} = \hat{j} \times \hat{j} = \hat{k} \times \hat{k} = 0,$$

$$\hat{i} \times \hat{j} = \hat{k}, \quad \hat{j} \times \hat{k} = \hat{i}, \quad \hat{k} \times \hat{i} = \hat{j}$$

Any vector \vec{a} with components a_x , a_y , and a_z along the x , y , and z axes can be written as

$$\vec{a} = a_x \hat{i} + a_y \hat{j} + a_z \hat{k}.$$

Let \vec{a} , \vec{b} , and \vec{c} be arbitrary vectors with magnitudes a , b , and c . Then

$$\vec{a} \times (\vec{b} + \vec{c}) = (\vec{a} \times \vec{b}) + (\vec{a} \times \vec{c})$$

$$(s\vec{a}) \times \vec{b} = \vec{a} \times (s\vec{b}) = s(\vec{a} \times \vec{b}) \quad (s = \text{a scalar}).$$

Let θ be the smaller of the two angles between \vec{a} and \vec{b} . Then

$$\vec{a} \cdot \vec{b} = \vec{b} \cdot \vec{a} = a_x b_x + a_y b_y + a_z b_z = ab \cos \theta$$

$$\vec{a} \times \vec{b} = -\vec{b} \times \vec{a} = \begin{vmatrix} \hat{i} & \hat{j} & \hat{k} \\ a_x & a_y & a_z \\ b_x & b_y & b_z \end{vmatrix}$$

$$\begin{aligned} &= \hat{i} \begin{vmatrix} a_y & a_z \\ b_y & b_z \end{vmatrix} - \hat{j} \begin{vmatrix} a_x & a_z \\ b_x & b_z \end{vmatrix} + \hat{k} \begin{vmatrix} a_x & a_y \\ b_x & b_y \end{vmatrix} \\ &= (a_y b_z - b_y a_z) \hat{i} + (a_z b_x - b_z a_x) \hat{j} \\ &\quad + (a_x b_y - b_x a_y) \hat{k} \\ |\vec{a} \times \vec{b}| &= ab \sin \theta \end{aligned}$$

$$\vec{a} \cdot (\vec{b} \times \vec{c}) = \vec{b} \cdot (\vec{c} \times \vec{a}) = \vec{c} \cdot (\vec{a} \times \vec{b})$$

$$\vec{a} \times (\vec{b} \times \vec{c}) = (\vec{a} \cdot \vec{c})\vec{b} - (\vec{a} \cdot \vec{b})\vec{c}$$

Derivatives and Integrals

In what follows, the letters u and v stand for any functions of x , and a and m are constants. To each of the indefinite integrals should be added an arbitrary constant of integration. The *Handbook of Chemistry and Physics* (CRC Press Inc.) gives a more extensive tabulation.

$$1. \frac{dx}{dx} = 1$$

$$2. \frac{d}{dx}(au) = a \frac{du}{dx}$$

$$3. \frac{d}{dx}(u + v) = \frac{du}{dx} + \frac{dv}{dx}$$

$$4. \frac{d}{dx}x^m = mx^{m-1}$$

$$5. \frac{d}{dx}\ln x = \frac{1}{x}$$

$$6. \frac{d}{dx}(uv) = u \frac{dv}{dx} + v \frac{du}{dx}$$

$$7. \frac{d}{dx}e^x = e^x$$

$$8. \frac{d}{dx}\sin x = \cos x$$

$$9. \frac{d}{dx}\cos x = -\sin x$$

$$10. \frac{d}{dx}\tan x = \sec^2 x$$

$$11. \frac{d}{dx}\cot x = -\csc^2 x$$

$$12. \frac{d}{dx}\sec x = \tan x \sec x$$

$$13. \frac{d}{dx}\csc x = -\cot x \csc x$$

$$14. \frac{d}{dx}e^u = e^u \frac{du}{dx}$$

$$15. \frac{d}{dx}\sin u = \cos u \frac{du}{dx}$$

$$16. \frac{d}{dx}\cos u = -\sin u \frac{du}{dx}$$

$$1. \int dx = x$$

$$2. \int au \, dx = a \int u \, dx$$

$$3. \int (u + v) \, dx = \int u \, dx + \int v \, dx$$

$$4. \int x^m \, dx = \frac{x^{m+1}}{m+1} \quad (m \neq -1)$$

$$5. \int \frac{dx}{x} = \ln|x|$$

$$6. \int u \frac{dv}{dx} \, dx = uv - \int v \frac{du}{dx} \, dx$$

$$7. \int e^x \, dx = e^x$$

$$8. \int \sin x \, dx = -\cos x$$

$$9. \int \cos x \, dx = \sin x$$

$$10. \int \tan x \, dx = \ln|\sec x|$$

$$11. \int \sin^2 x \, dx = \frac{1}{2}x - \frac{1}{4}\sin 2x$$

$$12. \int e^{-ax} \, dx = -\frac{1}{a}e^{-ax}$$

$$13. \int xe^{-ax} \, dx = -\frac{1}{a^2}(ax + 1)e^{-ax}$$

$$14. \int x^2 e^{-ax} \, dx = -\frac{1}{a^3}(a^2x^2 + 2ax + 2)e^{-ax}$$

$$15. \int_0^\infty x^n e^{-ax} \, dx = \frac{n!}{a^{n+1}}$$

$$16. \int_0^\infty x^{2n} e^{-ax^2} \, dx = \frac{1 \cdot 3 \cdot 5 \cdots (2n-1)}{2^{n+1}a^n} \sqrt{\frac{\pi}{a}}$$

$$17. \int \frac{dx}{\sqrt{x^2 + a^2}} = \ln(x + \sqrt{x^2 + a^2})$$

$$18. \int \frac{x \, dx}{(x^2 + a^2)^{3/2}} = -\frac{1}{(x^2 + a^2)^{1/2}}$$

$$19. \int \frac{dx}{(x^2 + a^2)^{3/2}} = \frac{x}{a^2(x^2 + a^2)^{1/2}}$$

$$20. \int_0^\infty x^{2n+1} e^{-ax^2} \, dx = \frac{n!}{2a^{n+1}} \quad (a > 0)$$

$$21. \int \frac{x \, dx}{x+d} = x - d \ln(x+d)$$

Properties of the Elements

All physical properties are for a pressure of 1 atm unless otherwise specified.

Element	Symbol	Atomic Number <i>Z</i>	Molar Mass, g/mol	Density, g/cm ³ at 20°C	Melting Point, °C	Boiling Point, °C	Specific Heat, J/(g · °C) at 25°C
Actinium	Ac	89	(227)	10.06	1323	(3473)	0.092
Aluminum	Al	13	26.9815	2.699	660	2450	0.900
Americium	Am	95	(243)	13.67	1541	—	—
Antimony	Sb	51	121.75	6.691	630.5	1380	0.205
Argon	Ar	18	39.948	1.6626×10^{-3}	-189.4	-185.8	0.523
Arsenic	As	33	74.9216	5.78	817 (28 atm)	613	0.331
Astatine	At	85	(210)	—	(302)	—	—
Barium	Ba	56	137.34	3.594	729	1640	0.205
Berkelium	Bk	97	(247)	14.79	—	—	—
Beryllium	Be	4	9.0122	1.848	1287	2770	1.83
Bismuth	Bi	83	208.980	9.747	271.37	1560	0.122
Bohrium	Bh	107	262.12	—	—	—	—
Boron	B	5	10.811	2.34	2030	—	1.11
Bromine	Br	35	79.909	3.12 (liquid)	-7.2	58	0.293
Cadmium	Cd	48	112.40	8.65	321.03	765	0.226
Calcium	Ca	20	40.08	1.55	838	1440	0.624
Californium	Cf	98	(251)	—	—	—	—
Carbon	C	6	12.01115	2.26	3727	4830	0.691
Cerium	Ce	58	140.12	6.768	804	3470	0.188
Cesium	Cs	55	132.905	1.873	28.40	690	0.243
Chlorine	Cl	17	35.453	3.214×10^{-3} (0°C)	-101	-34.7	0.486
Chromium	Cr	24	51.996	7.19	1857	2665	0.448
Cobalt	Co	27	58.9332	8.85	1495	2900	0.423
Copernicium	Cp	112	(285)	—	—	—	—
Copper	Cu	29	63.54	8.96	1083.40	2595	0.385
Curium	Cm	96	(247)	13.3	—	—	—
Darmstadtium	Ds	110	(271)	—	—	—	—
Dubnium	Db	105	262.114	—	—	—	—
Dysprosium	Dy	66	162.50	8.55	1409	2330	0.172
Einsteinium	Es	99	(254)	—	—	—	—
Erbium	Er	68	167.26	9.15	1522	2630	0.167
Europium	Eu	63	151.96	5.243	817	1490	0.163
Fermium	Fm	100	(237)	—	—	—	—
Fluorine	F	9	18.9984	1.696×10^{-3} (0°C)	-219.6	-188.2	0.753
Francium	Fr	87	(223)	—	(27)	—	—
Gadolinium	Gd	64	157.25	7.90	1312	2730	0.234
Gallium	Ga	31	69.72	5.907	29.75	2237	0.377
Germanium	Ge	32	72.59	5.323	937.25	2830	0.322
Gold	Au	79	196.967	19.32	1064.43	2970	0.131

Element	Symbol	Atomic Number Z	Molar Mass, g/mol	Density, g/cm ³ at 20°C	Melting Point, °C	Boiling Point, °C	Specific Heat, J/(g · °C) at 25°C
Hafnium	Hf	72	178.49	13.31	2227	5400	0.144
Hassium	Hs	108	(265)	—	—	—	—
Helium	He	2	4.0026	0.1664 × 10 ⁻³	-269.7	-268.9	5.23
Holmium	Ho	67	164.930	8.79	1470	2330	0.165
Hydrogen	H	1	1.00797	0.08375 × 10 ⁻³	-259.19	-252.7	14.4
Indium	In	49	114.82	7.31	156.634	2000	0.233
Iodine	I	53	126.9044	4.93	113.7	183	0.218
Iridium	Ir	77	192.2	22.5	2447	(5300)	0.130
Iron	Fe	26	55.847	7.874	1536.5	3000	0.447
Krypton	Kr	36	83.80	3.488 × 10 ⁻³	-157.37	-152	0.247
Lanthanum	La	57	138.91	6.189	920	3470	0.195
Lawrencium	Lr	103	(257)	—	—	—	—
Lead	Pb	82	207.19	11.35	327.45	1725	0.129
Lithium	Li	3	6.939	0.534	180.55	1300	3.58
Lutetium	Lu	71	174.97	9.849	1663	1930	0.155
Magnesium	Mg	12	24.312	1.738	650	1107	1.03
Manganese	Mn	25	54.9380	7.44	1244	2150	0.481
Meitnerium	Mt	109	(266)	—	—	—	—
Mendelevium	Md	101	(256)	—	—	—	—
Mercury	Hg	80	200.59	13.55	-38.87	357	0.138
Molybdenum	Mo	42	95.94	10.22	2617	5560	0.251
Neodymium	Nd	60	144.24	7.007	1016	3180	0.188
Neon	Ne	10	20.183	0.8387 × 10 ⁻³	-248.597	-246.0	1.03
Neptunium	Np	93	(237)	20.25	637	—	1.26
Nickel	Ni	28	58.71	8.902	1453	2730	0.444
Niobium	Nb	41	92.906	8.57	2468	4927	0.264
Nitrogen	N	7	14.0067	1.1649 × 10 ⁻³	-210	-195.8	1.03
Nobelium	No	102	(255)	—	—	—	—
Osmium	Os	76	190.2	22.59	3027	5500	0.130
Oxygen	O	8	15.9994	1.3318 × 10 ⁻³	-218.80	-183.0	0.913
Palladium	Pd	46	106.4	12.02	1552	3980	0.243
Phosphorus	P	15	30.9738	1.83	44.25	280	0.741
Platinum	Pt	78	195.09	21.45	1769	4530	0.134
Plutonium	Pu	94	(244)	19.8	640	3235	0.130
Polonium	Po	84	(210)	9.32	254	—	—
Potassium	K	19	39.102	0.862	63.20	760	0.758
Praseodymium	Pr	59	140.907	6.773	931	3020	0.197
Promethium	Pm	61	(145)	7.22	(1027)	—	—
Protactinium	Pa	91	(231)	15.37 (estimated)	(1230)	—	—
Radium	Ra	88	(226)	5.0	700	—	—
Radon	Rn	86	(222)	9.96 × 10 ⁻³ (0°C)	(-71)	-61.8	0.092
Rhenium	Re	75	186.2	21.02	3180	5900	0.134
Rhodium	Rh	45	102.905	12.41	1963	4500	0.243
Roentgenium	Rg	111	(280)	—	—	—	—
Rubidium	Rb	37	85.47	1.532	39.49	688	0.364
Ruthenium	Ru	44	101.107	12.37	2250	4900	0.239
Rutherfordium	Rf	104	261.11	—	—	—	—

A-14 APPENDIX F PROPERTIES OF THE ELEMENTS

Element	Symbol	Atomic Number Z	Molar Mass, g/mol	Density, g/cm ³ at 20°C	Melting Point, °C	Boiling Point, °C	Specific Heat, J/(g · °C) at 25°C
Samarium	Sm	62	150.35	7.52	1072	1630	0.197
Scandium	Sc	21	44.956	2.99	1539	2730	0.569
Seaborgium	Sg	106	263.118	—	—	—	—
Selenium	Se	34	78.96	4.79	221	685	0.318
Silicon	Si	14	28.086	2.33	1412	2680	0.712
Silver	Ag	47	107.870	10.49	960.8	2210	0.234
Sodium	Na	11	22.9898	0.9712	97.85	892	1.23
Strontium	Sr	38	87.62	2.54	768	1380	0.737
Sulfur	S	16	32.064	2.07	119.0	444.6	0.707
Tantalum	Ta	73	180.948	16.6	3014	5425	0.138
Technetium	Tc	43	(99)	11.46	2200	—	0.209
Tellurium	Te	52	127.60	6.24	449.5	990	0.201
Terbium	Tb	65	158.924	8.229	1357	2530	0.180
Thallium	Tl	81	204.37	11.85	304	1457	0.130
Thorium	Th	90	(232)	11.72	1755	(3850)	0.117
Thulium	Tm	69	168.934	9.32	1545	1720	0.159
Tin	Sn	50	118.69	7.2984	231.868	2270	0.226
Titanium	Ti	22	47.90	4.54	1670	3260	0.523
Tungsten	W	74	183.85	19.3	3380	5930	0.134
Unnamed	Uut	113	(284)	—	—	—	—
Unnamed	Unq	114	(289)	—	—	—	—
Unnamed	Uup	115	(288)	—	—	—	—
Unnamed	Uuh	116	(293)	—	—	—	—
Unnamed	Uus	117	—	—	—	—	—
Unnamed	Uuo	118	(294)	—	—	—	—
Uranium	U	92	(238)	18.95	1132	3818	0.117
Vanadium	V	23	50.942	6.11	1902	3400	0.490
Xenon	Xe	54	131.30	5.495 × 10 ⁻³	-111.79	-108	0.159
Ytterbium	Yb	70	173.04	6.965	824	1530	0.155
Yttrium	Y	39	88.905	4.469	1526	3030	0.297
Zinc	Zn	30	65.37	7.133	419.58	906	0.389
Zirconium	Zr	40	91.22	6.506	1852	3580	0.276

The values in parentheses in the column of molar masses are the mass numbers of the longest-lived isotopes of those elements that are radioactive. Melting points and boiling points in parentheses are uncertain.

The data for gases are valid only when these are in their usual molecular state, such as H₂, He, O₂, Ne, etc. The specific heats of the gases are the values at constant pressure.

Source: Adapted from J. Emsley, *The Elements*, 3rd ed., 1998, Clarendon Press, Oxford. See also www.webelements.com for the latest values and newest elements.

Periodic Table of the Elements

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5	Rb	37	Sr	38	Y	39	Zr	40	Nb	42	Mo	43	Tc	44	Ru	45	Rh	46	Pd	47	Ag	48	Cd	49	In	50	Ga	32	Si	14	P	15	S	16	Cl	17	Ar																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																																													
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Evidence for the discovery of elements 113 through 118 has been reported. See www.webelements.com for the latest information and newest elements.

CHAPTER 1

P **1.** (a) $4.00 \times 10^4 \text{ km}$; (b) $5.10 \times 10^8 \text{ km}^2$; (c) $1.08 \times 10^{12} \text{ km}^3$
3. (a) $10^9 \mu\text{m}$; (b) 10^{-4} ; (c) $9.1 \times 10^5 \mu\text{m}$ **5.** (a) 160 rods; (b) 40 chains **7.** $1.1 \times 10^3 \text{ acre-feet}$ **9.** $1.9 \times 10^{22} \text{ cm}^3$ **11.** (a) 1.43; (b) 0.864 **13.** (a) 495 s; (b) 141 s; (c) 198 s; (d) -245 s **15.** $1.21 \times 10^{12} \mu\text{s}$ **17.** C, D, A, B, E; the important criterion is the consistency of the daily variation, not its magnitude **19.** $5.2 \times 10^6 \text{ m}$ **21.** 9.0 $\times 10^{49}$ atoms **23.** (a) $1 \times 10^3 \text{ kg}$; (b) 158 kg/s **25.** $1.9 \times 10^5 \text{ kg}$ **27.** (a) $1.18 \times 10^{-29} \text{ m}^3$; (b) 0.282 nm **29.** $1.75 \times 10^3 \text{ kg}$ **31.** 1.43 kg/min **33.** (a) 293 U.S. bushels; (b) $3.81 \times 10^3 \text{ U.S. bushels}$ **35.** (a) 22 pecks; (b) 5.5 Imperial bushels; (c) 200 L **37.** $8 \times 10^2 \text{ km}$ **39.** (a) 18.8 gallons; (b) 22.5 gallons **41.** 0.3 cord **43.** 3.8 mg/s **45.** (a) yes; (b) 8.6 universe seconds **47.** 0.12 AU/min **49.** (a) 3.88; (b) 7.65; (c) 156 ken³; (d) $1.19 \times 10^3 \text{ m}^3$ **51.** (a) 3.9 m, 4.8 m; (b) $3.9 \times 10^3 \text{ mm}$, $4.8 \times 10^3 \text{ mm}$; (c) 2.2 m^3 , 4.2 m^3 **53.** (a) $4.9 \times 10^{-6} \text{ pc}$; (b) $1.6 \times 10^{-5} \text{ ly}$

CHAPTER 2

CP **1.** b and c **2.** (check the derivative dx/dt) (a) 1 and 4; (b) 2 and 3 **3.** (a) plus; (b) minus; (c) minus; (d) plus **4.** 1 and 4 ($a = d^2x/dt^2$ must be constant) **5.** (a) plus (upward displacement on y axis); (b) minus (downward displacement on y axis); (c) $a = -g = -9.8 \text{ m/s}^2$

Q **1.** (a) negative; (b) positive; (c) yes; (d) positive; (e) constant **3.** (a) all tie; (b) 4, tie of 1 and 2, then 3 **5.** (a) positive direction; (b) negative direction; (c) 3 and 5; (d) 2 and 6 tie, then 3 and 5 tie, then 1 and 4 tie (zero) **7.** (a) D; (b) E **9.** (a) 3, 2, 1; (b) 1, 2, 3; (c) all tie; (d) 1, 2, 3

P **1.** 13 m **3.** (a) +40 km/h; (b) 40 km/h **5.** (a) 0; (b) -2 m; (c) 0; (d) 12 m; (e) +12 m; (f) +7 m/s **7.** 60 km **9.** 1.4 m **11.** 128 km/h **13.** (a) 73 km/h; (b) 68 km/h; (c) 70 km/h; (d) 0 **15.** (a) -6 m/s; (b) -x direction; (c) 6 m/s; (d) decreasing; (e) 2 s; (f) no **17.** (a) 28.5 cm/s; (b) 18.0 cm/s; (c) 40.5 cm/s; (d) 28.1 cm/s; (e) 30.3 cm/s **19.** -20 m/s² **21.** (a) 1.10 m/s; (b) 6.11 mm/s²; (c) 1.47 m/s; (d) 6.11 mm/s² **23.** $1.62 \times 10^{15} \text{ m/s}^2$ **25.** (a) 30 s; (b) 300 m **27.** (a) +1.6 m/s; (b) +18 m/s **29.** (a) 10.6 m; (b) 41.5 s **31.** (a) $3.1 \times 10^6 \text{ s}$; (b) $4.6 \times 10^{13} \text{ m}$ **33.** (a) 3.56 m/s^2 ; (b) 8.43 m/s **35.** 0.90 m/s^2 **37.** (a) 4.0 m/s^2 ; (b) +x **39.** (a) -2.5 m/s²; (b) 1; (d) 0; (e) 2 **41.** 40 m **43.** (a) 0.994 m/s^2 **45.** (a) 31 m/s; (b) 6.4 s **47.** (a) 29.4 m; (b) 2.45 s **49.** (a) 5.4 s; (b) 41 m/s **51.** (a) 20 m; (b) 59 m **53.** 4.0 m/s **55.** (a) 857 m/s²; (b) up **57.** (a) $1.26 \times 10^3 \text{ m/s}^2$; (b) up **59.** (a) 89 cm; (b) 22 cm **61.** 20.4 m **63.** 2.34 m **65.** (a) 2.25 m/s; (b) 3.90 m/s **67.** 0.56 m/s **69.** 100 m **71.** (a) 2.00 s; (b) 12 cm; (c) -9.00 cm/s²; (d) right; (e) left; (f) 3.46 s **73.** (a) 82 m; (b) 19 m/s **75.** (a) 0.74 s; (b) 6.2 m/s² **77.** (a) 3.1 m/s^2 ; (b) 45 m; (c) 13 s **79.** 17 m/s **81.** +47 m/s **83.** (a) 1.23 cm; (b) 4 times; (c) 9 times; (d) 16 times; (e) 25 times **85.** 25 km/h **87.** 1.2 h **89.** 4H **91.** (a) 3.2 s; (b) 1.3 s **93.** (a) 8.85 m/s; (b) 1.00 m **95.** (a) 2.0 m/s^2 ; (b) 12 m/s; (c) 45 m **97.** (a) 48.5 m/s; (b) 4.95 s; (c) 34.3 m/s; (d) 3.50 s **99.** 22.0 m/s **101.** (a) $v = (v_0^2 + 2gh)^{0.5}$; (b) $t = [(v_0^2 + 2gh)^{0.5} - v_0]/g$; (c) same as (a); (d) $t = [(v_0^2 + 2gh)^{0.5} + v_0]/g$, greater

CHAPTER 3

CP **1.** (a) 7 m (\vec{a} and \vec{b} are in same direction); (b) 1 m (\vec{a} and \vec{b} are in opposite directions) **2.** c, d, f (components must be head to tail; \vec{a} must extend from tail of one component to head of the other) **3.** (a) +, +; (b) +, -; (c) +, + (draw vector from tail of \vec{d}_1 to head of \vec{d}_2)

4. (a) 90° ; (b) 0° (vectors are parallel—same direction); (c) 180° (vectors are antiparallel—opposite directions) **5.** (a) 0° or 180° ; (b) 90°

Q **1.** yes, when the vectors are in same direction **3.** Either the sequence \vec{d}_2, \vec{d}_1 or the sequence $\vec{d}_2, \vec{d}_2, \vec{d}_3$ **5.** all but (e) **7.** (a) yes; (b) yes; (c) no **9.** (a) +x for (1), +z for (2), +z for (3); (b) -x for (1), -z for (2), -z for (3)

P **1.** (a) -2.5 m; (b) -6.9 m **3.** (a) 47.2 m; (b) 122° **5.** (a) 156 km; (b) 39.8° west of due north **7.** (a) 6.42 m; (b) no; (c) yes; (d) yes; (e) a possible answer: $(4.30 \text{ m})\hat{i} + (3.70 \text{ m})\hat{j} + (3.00 \text{ m})\hat{k}$; (f) 7.96 m **9.** (a) $(3.0 \text{ m})\hat{i} - (2.0 \text{ m})\hat{j} + (5.0 \text{ m})\hat{k}$; (b) $(5.0 \text{ m})\hat{i} - (4.0 \text{ m})\hat{j} - (3.0 \text{ m})\hat{k}$; (c) $(-5.0 \text{ m})\hat{i} + (4.0 \text{ m})\hat{j} + (3.0 \text{ m})\hat{k}$ **11.** (a) $(-9.0 \text{ m})\hat{i} + (10 \text{ m})\hat{j}$; (b) 13 m; (c) 132° **13.** 4.74 km **15.** (a) 1.59 m; (b) 12.1 m; (c) 12.2 m; (d) 82.5° **17.** (a) 38 m; (b) -37.5°; (c) 130 m; (d) 1.2° ; (e) 62 m; (f) 130° **19.** 5.39 m at 21.8° left of forward **21.** (a) -70.0 cm; (b) 80.0 cm; (c) 141 cm; (d) -172° **23.** 3.2 **25.** 2.6 km **27.** (a) $8\hat{i} + 16\hat{j}$; (b) $2\hat{i} + 4\hat{j}$ **29.** (a) 7.5 cm; (b) 90° ; (c) 8.6 cm; (d) 48° **31.** (a) $\hat{a}\hat{i} + \hat{a}\hat{j} + \hat{a}\hat{k}$; (b) $-\hat{a}\hat{i} + \hat{a}\hat{j} + \hat{a}\hat{k}$; (c) $\hat{a}\hat{i} - \hat{a}\hat{j} + \hat{a}\hat{k}$; (d) $-\hat{a}\hat{i} - \hat{a}\hat{j} + \hat{a}\hat{k}$; (e) 54.7° ; (f) 30.5° **33.** (a) 12; (b) +z; (c) 12; (d) -z; (e) 12; (f) +z **35.** (a) -18.8 units; (b) 26.9 units, +z direction **37.** (a) -21; (b) -9; (c) $5\hat{i} - 11\hat{j} - 9\hat{k}$ **39.** 70.5° **41.** 22° **43.** (a) 3.00 m; (b) 0; (c) 3.46 m; (d) 2.00 m; (e) -5.00 m; (f) 8.66 m; (g) -6.67; (h) 4.33 **45.** (a) -83.4; (b) $(1.14 \times 10^3)\hat{k}$; (c) 1.14×10^3 , θ not defined, $\phi = 0^\circ$; (d) 90.0° ; (e) $-5.14\hat{i} + 6.13\hat{j} + 3.00\hat{k}$; (f) 8.54, $\theta = 130^\circ$, $\phi = 69.4^\circ$ **47.** (a) 140° ; (b) 90.0° ; (c) 99.1° **49.** (a) 103 km; (b) 60.9° north of due west **51.** (a) 27.8 m; (b) 13.4 m **53.** (a) 30; (b) 52 **55.** (a) -2.83 m; (b) -2.83 m; (c) 5.00 m; (d) 0; (e) 3.00 m; (f) 5.20 m; (g) 5.17 m; (h) 2.37 m; (i) 5.69 m; (j) 25° north of due east; (k) 5.69 m; (l) 25° south of due west **57.** 4.1 **59.** (a) $(9.19 \text{ m})\hat{i} + (7.71 \text{ m})\hat{j}$; (b) $(14.0 \text{ m})\hat{i}' + (3.41 \text{ m})\hat{j}'$ **61.** (a) $11\hat{i} + 5.0\hat{j} - 7.0\hat{k}$; (b) 120° ; (c) -4.9; (d) 7.3 **63.** (a) 3.0 m^2 ; (b) 52 m³; (c) $(11 \text{ m}^2)\hat{i} + (9.0 \text{ m}^2)\hat{j} + (3.0 \text{ m}^2)\hat{k}$ **65.** (a) $(-40\hat{i} - 20\hat{j} + 25\hat{k}) \text{ m}$; (b) 45 m

CHAPTER 4

CP **1.** (draw \vec{v} tangent to path, tail on path) (a) first; (b) third **2.** (take second derivative with respect to time) (1) and (3) a_x and a_y are both constant and thus \vec{a} is constant; (2) and (4) a_y is constant but a_x is not, thus \vec{a} is not **3.** yes **4.** (a) v_x constant; (b) v_y initially positive, decreases to zero, and then becomes progressively more negative; (c) $a_x = 0$ throughout; (d) $a_y = -g$ throughout **5.** (a) $(-4 \text{ m/s})\hat{i}$; (b) $(-8 \text{ m/s}^2)\hat{i}$

Q **1.** a and c tie, then b **3.** decreases **5.** a, b, c **7.** (a) 0; (b) 350 km/h; (c) 350 km/h; (d) same (nothing changed about the vertical motion) **9.** (a) all tie; (b) all tie; (c) 3, 2, 1; (d) 3, 2, 1 **11.** 2, then 1 and 4 tie, then 3 **13.** (a) yes; (b) no; (c) yes

P **1.** (a) 6.2 m **3.** $(-2.0 \text{ m})\hat{i} + (6.0 \text{ m})\hat{j} - (10 \text{ m})\hat{k}$ **5.** (a) 7.59 km/h; (b) 22.5° east of due north **7.** $(-0.70 \text{ m/s})\hat{i} + (1.4 \text{ m/s})\hat{j} - (0.40 \text{ m/s})\hat{k}$ **9.** (a) 0.83 cm/s; (b) 0° ; (c) 0.11 m/s; (d) -63° **11.** (a) $(6.00 \text{ m})\hat{i} - (106 \text{ m})\hat{j}$; (b) $(19.0 \text{ m/s})\hat{i} - (224 \text{ m/s})\hat{j}$; (c) $(24.0 \text{ m/s}^2)\hat{i} - (336 \text{ m/s}^2)\hat{j}$; (d) -85.2° **13.** (a) $(8 \text{ m/s}^2)\hat{t}\hat{j} + (1 \text{ m/s})\hat{k}$; (b) $(8 \text{ m/s}^2)\hat{j}$ **15.** (a) $(-1.50 \text{ m/s})\hat{j}$; (b) $(4.50 \text{ m})\hat{i} - (2.25 \text{ m})\hat{j}$ **17.** $(32 \text{ m/s})\hat{i}$ **19.** (a) $(72.0 \text{ m})\hat{i} + (90.7 \text{ m})\hat{j}$; (b) 49.5° **21.** (a) 18 cm; (b) 1.9 m **23.** (a) 3.03 s; (b) 758 m; (c) 29.7 m/s **25.** 43.1 m/s (155 km/h) **27.** (a) 10.0 s; (b) 897 m **29.** 78.5° **31.** 3.35 m **33.** (a) 202 m/s; (b) 806 m; (c) 161 m/s; (d) -171 m/s **35.** 4.84 cm **37.** (a) 1.60 m; (b) 6.86 m; (c) 2.86 m **39.** (a) 32.3 m; (b) 21.9 m/s; (c) 40.4°; (d) below **41.** 55.5° **43.** (a) 11 m; (b) 23 m; (c) 17 m/s; (d) 63°

AN-2 ANSWERS

- 45.** (a) ramp; (b) 5.82 m; (c) 31.0° **47.** (a) yes; (b) 2.56 m **49.** (a) 31° ; (b) 63° **51.** (a) 2.3° ; (b) 1.4 m; (c) 18° **53.** (a) 75.0 m; (b) 31.9 m/s; (c) 66.9°; (d) 25.5 m **55.** the third **57.** (a) 7.32 m; (b) west; (c) north **59.** (a) 12 s; (b) 4.1 m/s^2 ; (c) down; (d) 4.1 m/s^2 ; (e) up **61.** (a) $1.3 \times 10^5 \text{ m/s}$; (b) $7.9 \times 10^5 \text{ m/s}^2$; (c) increase **63.** 2.92 m **65.** $(3.00 \text{ m/s}^2)\hat{i} + (6.00 \text{ m/s}^2)\hat{j}$ **67.** 160 m/s **69.** (a) 13 m/s^2 ; (b) eastward; (c) 13 m/s^2 ; (d) eastward **71.** 1.67 **73.** (a) $(80 \text{ km/h})\hat{i} - (60 \text{ km/h})\hat{j}$; (b) 0° ; (c) answers do not change **75.** 32 m/s **77.** 60° **79.** (a) 38 knots; (b) 1.5° east of due north; (c) 4.2 h; (d) 1.5° west of due south **81.** (a) $(-32 \text{ km/h})\hat{i} - (46 \text{ km/h})\hat{j}$; (b) $[(2.5 \text{ km}) - (32 \text{ km/h})t]\hat{i} + [(4.0 \text{ km}) - (46 \text{ km/h})t]\hat{j}$; (c) 0.084 h; (d) $2 \times 10^2 \text{ m}$ **83.** (a) -30° ; (b) 69 min; (c) 80 min; (d) 80 min; (e) 0° ; (f) 60 min **85.** (a) 2.7 km; (b) 76° clockwise **87.** (a) 44 m; (b) 13 m; (c) 8.9 m **89.** (a) 45 m; (b) 22 m/s **91.** (a) $2.6 \times 10^2 \text{ m/s}$; (b) 45 s; (c) increase **93.** (a) 63 km; (b) 18° south of due east; (c) 0.70 km/h; (d) 18° south of due east; (e) 1.6 km/h; (f) 1.2 km/h; (g) 33° north of due east **95.** (a) 1.5; (b) (36 m, 54 m) **97.** (a) 62 ms; (b) $4.8 \times 10^2 \text{ m/s}$ **99.** 2.64 m **101.** (a) 2.5 m; (b) 0.82 m; (c) 9.8 m/s^2 ; (d) 9.8 m/s^2 **103.** (a) 6.79 km/h; (b) 6.96° **105.** (a) 16 m/s; (b) 23° ; (c) above; (d) 27 m/s; (e) 57° ; (f) below **107.** (a) 4.2 m, 45° ; (b) 5.5 m, 68° ; (c) 6.0 m, 90° ; (d) 4.2 m, 135° ; (e) 0.85 m/s, 135° ; (f) 0.94 m/s, 90° ; (g) 0.94 m/s, 180° ; (h) 0.30 m/s^2 , 180° ; (i) 0.30 m/s^2 , 270° **109.** (a) $5.4 \times 10^{-13} \text{ m}$; (b) decrease **111.** (a) 0.034 m/s^2 ; (b) 84 min **113.** (a) 8.43 m; (b) -129° **115.** (a) 2.00 ns; (b) 2.00 mm; (c) $1.00 \times 10^7 \text{ m/s}$; (d) $2.00 \times 10^6 \text{ m/s}$ **117.** (a) 24 m/s; (b) 65° **119.** 93° from the car's direction of motion

CHAPTER 5

CP 1. c, d, and e (\vec{F}_1 and \vec{F}_2 must be head to tail, \vec{F}_{net} must be from tail of one of them to head of the other) **2.** (a) and (b) 2 N, leftward (acceleration is zero in each situation) **3.** (a) equal; (b) greater (acceleration is upward, thus net force on body must be upward) **4.** (a) equal; (b) greater; (c) less **5.** (a) increase; (b) yes; (c) same; (d) yes

Q 1. (a) 2, 3, 4; (b) 1, 3, 4; (c) 1, +y; 2, +x; 3, fourth quadrant; 4, third quadrant **3.** increase **5.** (a) 2 and 4; (b) 2 and 4 **7.** (a) M; (b) M'; (c) M'; (d) 2M; (e) 3M **9.** (a) 20 kg; (b) 18 kg; (c) 10 kg; (d) all tie; (e) 3, 2, 1 **11.** (a) increases from initial value mg ; (b) decreases from mg to zero (after which the block moves up away from the floor)

- P** 1. 2.9 m/s^2 **3.** (a) 1.88 N; (b) 0.684 N; (c) $(1.88 \text{ N})\hat{i} + (0.684 \text{ N})\hat{j}$ **5.** (a) $(0.86 \text{ m/s}^2)\hat{i} - (0.16 \text{ m/s}^2)\hat{j}$; (b) 0.88 m/s^2 ; (c) -11° **7.** (a) $(-32.0 \text{ N})\hat{i} - (20.8 \text{ N})\hat{j}$; (b) 38.2 N; (c) -147° **9.** (a) 8.37 N; (b) -133° ; (c) -125° **11.** 9.0 m/s **13.** (a) 4.0 kg; (b) 1.0 kg; (c) 4.0 kg; (d) 1.0 kg **15.** (a) 108 N; (b) 108 N; (c) 108 N **17.** (a) 42 N; (b) 72 N; (c) 4.9 m/s^2 **19.** $1.2 \times 10^5 \text{ N}$ **21.** (a) 11.7 N; (b) -59.0° **23.** (a) $(285 \text{ N})\hat{i} + (705 \text{ N})\hat{j}$; (b) $(285 \text{ N})\hat{i} - (115 \text{ N})\hat{j}$; (c) 307 N; (d) -22.0° ; (e) 3.67 m/s^2 ; (f) -22.0° **25.** (a) 0.022 m/s^2 ; (b) $8.3 \times 10^4 \text{ km}$; (c) $1.9 \times 10^3 \text{ m/s}$ **27.** 1.5 mm **29.** (a) 494 N; (b) up; (c) 494 N; (d) down **31.** (a) 1.18 m; (b) 0.674 s; (c) 3.50 m/s **33.** $1.8 \times 10^4 \text{ N}$ **35.** (a) 46.7° ; (b) 28.0° **37.** (a) 0.62 m/s^2 ; (b) 0.13 m/s^2 ; (c) 2.6 m **39.** (a) $2.2 \times 10^{-3} \text{ N}$; (b) $3.7 \times 10^{-3} \text{ N}$ **41.** (a) 1.4 m/s^2 ; (b) 4.1 m/s **43.** (a) 1.23 N; (b) 2.46 N; (c) 3.69 N; (d) 4.92 N; (e) 6.15 N; (f) 0.250 N **45.** (a) 31.3 kN; (b) 24.3 kN **47.** $6.4 \times 10^3 \text{ N}$ **49.** (a) 2.18 m/s^2 ; (b) 116 N; (c) 21.0 m/s^2 **51.** (a) 3.6 m/s^2 ; (b) 17 N **53.** (a) 0.970 m/s^2 ; (b) 11.6 N; (c) 34.9 N **55.** (a) 1.1 N **57.** (a) 0.735 m/s²; (b) down; (c) 20.8 N **59.** (a) 4.9 m/s^2 ; (b) 2.0 m/s^2 ; (c) up; (d) 120 N **61.** $2Ma/(a + g)$ **63.** (a) 8.0 m/s; (b) $+x$ **65.** (a) 0.653 m/s^3 ; (b) 0.896 m/s^3 ; (c) 6.50 s **67.** 81.7 N **69.** 2.4 N **71.** 16 N **73.** (a) 2.6 N; (b) 17° **75.** (a) 0; (b) 0.83 m/s²; (c) 0 **77.** (a) 0.74 m/s²; (b) 7.3 m/s² **79.** (a) 11 N; (b) 2.2 kg; (c) 0; (d) 2.2 kg **81.** 195 N **83.** (a) 4.6 m/s^2 ; (b) 2.6 m/s^2 **85.** (a) rope breaks; (b) 1.6 m/s^2 **87.** (a) 65 N; (b) 49 N **89.** (a) $4.6 \times 10^3 \text{ N}$; (b) $5.8 \times 10^3 \text{ N}$ **91.** (a) $1.8 \times 10^2 \text{ N}$; (b) $6.4 \times$

- 10^2 N **93.** (a) 44 N; (b) 78 N; (c) 54 N; (d) 152 N **95.** (a) 4 kg; (b) 6.5 m/s^2 ; (c) 13 N

CHAPTER 6

CP 1. (a) zero (because there is no attempt at sliding); (b) 5 N; (c) no; (d) yes; (e) 8 N **2.** (\vec{a}) is directed toward center of circular path) (a) \vec{a} downward, \vec{F}_N upward; (b) \vec{a} and \vec{F}_N upward

Q 1. (a) decrease; (b) decrease; (c) increase; (d) increase; (e) increase **3.** (a) same; (b) increases; (c) increases; (d) no **5.** (a) upward; (b) horizontal, toward you; (c) no change; (d) increases; (e) increases **7.** At first, \vec{f}_s is directed up the ramp and its magnitude increases from $mg \sin \theta$ until it reaches $f_{s,\max}$. Thereafter the force is kinetic friction directed up the ramp, with magnitude f_k (a constant value smaller than $f_{s,\max}$). **9.** 4, 3, then 1, 2, and 5 tie **11.** (a) all tie; (b) all tie; (c) 2, 3, 1

- P** 1. 36 m **3.** (a) $2.0 \times 10^2 \text{ N}$; (b) $1.2 \times 10^2 \text{ N}$ **5.** (a) 6.0 N; (b) 3.6 N ; (c) 3.1 N **7.** (a) $1.9 \times 10^2 \text{ N}$; (b) 0.56 m/s^2 **9.** (a) 11 N; (b) 0.14 m/s^2 **11.** (a) $3.0 \times 10^2 \text{ N}$; (b) 1.3 m/s^2 **13.** (a) $1.3 \times 10^2 \text{ N}$; (b) no; (c) $1.1 \times 10^2 \text{ N}$; (d) 46 N; (e) 17 N **15.** 2° **17.** (a) $(17 \text{ N})\hat{i}$; (b) $(20 \text{ N})\hat{i}$; (c) $(15 \text{ N})\hat{i} + (5.0 \text{ N})\hat{j}$ **21.** (a) 19° ; (b) 3.3 kN **23.** 0.37 **25.** $1.0 \times 10^2 \text{ N}$ **27.** (a) 0; (b) $(-3.9 \text{ m/s}^2)\hat{i}$; (c) $(-1.0 \text{ m/s}^2)\hat{i}$ **29.** (a) 66 N; (b) 2.3 m/s^2 **31.** (a) 3.5 m/s^2 ; (b) 0.21 N **33.** 9.9 s **35.** $4.9 \times 10^2 \text{ N}$ **37.** (a) $3.2 \times 10^2 \text{ km/h}$; (b) $6.5 \times 10^2 \text{ km/h}$; (c) no **39.** 2.3 **41.** 0.60 **43.** 21 m **45.** (a) light; (b) 778 N; (c) 223 N; (d) 1.11 kN **47.** (a) 10 s; (b) $4.9 \times 10^2 \text{ N}$; (c) $1.1 \times 10^3 \text{ N}$ **49.** $1.37 \times 10^3 \text{ N}$ **51.** 2.2 km **53.** 12° **55.** $2.6 \times 10^3 \text{ N}$ **57.** 1.81 m/s **59.** (a) 8.74 N; (b) 37.9 N; (c) 6.45 m/s; (d) radially inward **61.** (a) 27 N; (b) 3.0 m/s^2 **63.** (b) 240 N; (c) 0.60 **65.** (a) 69 km/h; (b) 139 km/h; (c) yes **67.** $g(\sin \theta - 2^{0.5} \mu_k \cos \theta)$ **69.** 3.4 m/s² **71.** (a) 35.3 N; (b) 39.7 N; (c) 320 N **73.** (a) 7.5 m/s^2 ; (b) down; (c) 9.5 m/s^2 ; (d) down **75.** (a) $3.0 \times 10^5 \text{ N}$; (b) 1.2° **77.** 147 m/s **79.** (a) 13 N; (b) 1.6 m/s^2 **81.** (a) 275 N; (b) 877 N **83.** (a) 84.2 N; (b) 52.8 N; (c) 1.87 m/s^2 **85.** 3.4% **87.** (a) $3.21 \times 10^3 \text{ N}$; (b) yes **89.** (a) 222 N; (b) 334 N; (c) 311 N; (d) 311 N; (e) c, d **91.** (a) $v_0^2/(4g \sin \theta)$; (b) no **93.** (a) 0.34; (b) 0.24 **95.** (a) $\mu_k mg/(\sin \theta - \mu_k \cos \theta)$; (b) $\theta_0 = \tan^{-1} \mu_s$ **97.** 0.18

CHAPTER 7

CP 1. (a) decrease; (b) same; (c) negative, zero **2.** (a) positive; (b) negative; (c) zero **3.** zero

Q 1. all tie **3.** (a) positive; (b) negative; (c) negative **5.** b (positive work), a (zero work), c (negative work), d (more negative work) **7.** all tie **9.** (a) A; (b) B

- P** 1. (a) $2.9 \times 10^7 \text{ m/s}$; (b) $2.1 \times 10^{-13} \text{ J}$ **3.** (a) $5 \times 10^{14} \text{ J}$; (b) 0.1 megaton TNT; (c) 8 bombs **5.** (a) 2.4 m/s ; (b) 4.8 m/s **7.** 0.96 J **9.** 20 J **11.** (a) 62.3° ; (b) 118° **13.** (a) $1.7 \times 10^2 \text{ N}$; (b) $3.4 \times 10^2 \text{ m}$; (c) $-5.8 \times 10^4 \text{ J}$; (d) $3.4 \times 10^2 \text{ N}$; (e) $1.7 \times 10^2 \text{ m}$; (f) $-5.8 \times 10^4 \text{ J}$ **15.** (a) 1.50 J; (b) increases **17.** (a) 12 kJ; (b) -11 kJ ; (c) 1.1 kJ ; (d) 5.4 m/s **19.** 25 J **21.** (a) $-3Mgd/4$; (b) Mgd ; (c) $Mgd/4$; (d) $(gd/2)^{0.5}$ **23.** 4.41 J **25.** (a) 25.9 kJ; (b) 2.45 N **27.** (a) 7.2 J; (b) 7.2 J; (c) 0; (d) -25 J **29.** (a) 0.90 J; (b) 2.1 J; (c) 0 **31.** (a) 6.6 m/s; (b) 4.7 m **33.** (a) 0.12 m; (b) 0.36 J; (c) -0.36 J ; (d) 0.060 m; (e) 0.090 J **35.** (a) 0; (b) 0 **37.** (a) 42 J; (b) 30 J; (c) 12 J; (d) 6.5 m/s, +x axis; (e) 5.5 m/s, +x axis; (f) 3.5 m/s, +x axis **39.** 4.00 N/m **41.** $5.3 \times 10^2 \text{ J}$ **43.** (a) 0.83 J; (b) 2.5 J; (c) 4.2 J; (d) 5.0 W **45.** $4.9 \times 10^2 \text{ W}$ **47.** (a) $1.0 \times 10^2 \text{ J}$; (b) 8.4 W **49.** $7.4 \times 10^2 \text{ W}$ **51.** (a) 32.0 J; (b) 8.00 W; (c) 78.2° **53.** (a) 1.20 J; (b) 1.10 m/s **55.** (a) $1.8 \times 10^5 \text{ ft} \cdot \text{lb}$; (b) 0.55 hp **57.** (a) 797 N; (b) 0; (c) -1.55 kJ ; (d) 0; (e) 1.55 kJ; (f) F varies during displacement **59.** (a) $1 \times 10^5 \text{ megatons TNT}$; (b) $1 \times 10^7 \text{ bombs}$ **61.** -6 J **63.** (a) 314 J; (b) -155 J ; (c) 0; (d) 158 J **65.** (a) 98 N; (b) 4.0 cm; (c) 3.9 J; (d) -3.9 J **67.** (a) 23 mm; (b) 45 N **69.** 165 kW

- 71.** –37 J **73.** (a) 13 J; (b) 13 J **75.** 235 kW **77.** (a) 6 J; (b) 6.0 J
79. (a) 0.6 J; (b) 0; (c) –0.6 J

CHAPTER 8

- CP** **1.** no (consider round trip on the small loop) **2.** 3, 1, 2 (see Eq. 8-6) **3.** (a) all tie; (b) all tie **4.** (a) $CD, AB, BC(0)$ (check slope magnitudes); (b) positive direction of x **5.** all tie
Q **1.** (a) 3, 2, 1; (b) 1, 2, 3 **3.** (a) 12 J; (b) –2 J **5.** (a) increasing; (b) decreasing; (c) decreasing; (d) constant in AB and BC , decreasing in CD **7.** +30 J **9.** 2, 1, 3
P **1.** 89 N/cm **3.** (a) 167 J; (b) –167 J; (c) 196 J; (d) 29 J; (e) 167 J; (f) –167 J; (g) 296 J; (h) 129 J **5.** (a) 4.31 mJ; (b) –4.31 mJ; (c) 4.31 mJ; (d) –4.31 mJ; (e) all increase **7.** (a) 13.1 J; (b) –13.1 J; (c) 13.1 J; (d) all increase **9.** (a) 17.0 m/s; (b) 26.5 m/s; (c) 33.4 m/s; (d) 56.7 m; (e) all the same **11.** (a) 2.08 m/s; (b) 2.08 m/s; (c) increase **13.** (a) 0.98 J; (b) –0.98 J; (c) 3.1 N/cm **15.** (a) 2.6×10^2 m; (b) same; (c) decrease **17.** (a) 2.5 N; (b) 0.31 N; (c) 30 cm **19.** (a) 784 N/m; (b) 62.7 J; (c) 62.7 J; (d) 80.0 cm **21.** (a) 8.35 m/s; (b) 4.33 m/s; (c) 7.45 m/s; (d) both decrease **23.** (a) 4.85 m/s; (b) 2.42 m/s **25.** -3.2×10^2 J **27.** (a) no; (b) 9.3×10^2 N **29.** (a) 35 cm; (b) 1.7 m/s **31.** (a) 39.2 J; (b) 39.2 J; (c) 4.00 m **33.** (a) 2.40 m/s; (b) 4.19 m/s **35.** (a) 39.6 cm; (b) 3.64 cm **37.** –18 mJ **39.** (a) 2.1 m/s; (b) 10 N; (c) + x direction; (d) 5.7 m; (e) 30 N; (f) – x direction **41.** (a) –3.7 J; (c) 1.3 m; (d) 9.1 m; (e) 2.2 J; (f) 4.0 m; (g) $(4 - x)e^{-x/4}$; (h) 4.0 m **43.** (a) 5.6 J; (b) 3.5 J **45.** (a) 30.1 J; (b) 30.1 J; (c) 0.225 **47.** 0.53 J **49.** (a) –2.9 kJ; (b) 3.9×10^2 J; (c) 2.1×10^2 N **51.** (a) 1.5 MJ; (b) 0.51 MJ; (c) 1.0 MJ; (d) 63 m/s **53.** (a) 67 J; (b) 67 J; (c) 46 cm **55.** (a) –0.90 J; (b) 0.46 J; (c) 1.0 m/s **57.** 1.2 m **59.** (a) 19.4 m; (b) 19.0 m/s **61.** (a) 1.5×10^{-2} N; (b) (3.8×10^2) g **63.** (a) 7.4 m/s; (b) 90 cm; (c) 2.8 m; (d) 15 m **65.** 20 cm **67.** (a) 7.0 J; (b) 22 J **69.** 3.7 J **71.** 4.33 m/s **73.** 25 J **75.** (a) 4.9 m/s; (b) 4.5 N; (c) 71°; (d) same **77.** (a) 4.8 N; (b) + x direction; (c) 1.5 m; (d) 13.5 m; (e) 3.5 m/s **79.** (a) 24 kJ; (b) 4.7×10^2 N **81.** (a) 5.00 J; (b) 9.00 J; (c) 11.0 J; (d) 3.00 J; (e) 12.0 J; (f) 2.00 J; (g) 13.0 J; (h) 1.00 J; (i) 13.0 J; (j) 1.00 J; (l) 11.0 J; (m) 10.8 m; (n) It returns to $x = 0$ and stops. **83.** (a) 6.0 kJ; (b) 6.0×10^2 W; (c) 3.0×10^2 W; (d) 9.0 $\times 10^2$ W **85.** 880 MW **87.** (a) $v_0 = (2gL)^{0.5}$; (b) 5mg; (c) $-mgL$; (d) $-2mgL$ **89.** (a) 109 J; (b) 60.3 J; (c) 68.2 J; (d) 41.0 J **91.** (a) 2.7 J; (b) 1.8 J; (c) 0.39 m **93.** (a) 10 m; (b) 49 N; (c) 4.1 m; (d) 1.2×10^2 N **95.** (a) 5.5 m/s; (b) 5.4 m; (c) same **97.** 80 mJ **99.** 24 W **101.** –12 J **103.** (a) 8.8 m/s; (b) 2.6 kJ; (c) 1.6 kW **105.** (a) 7.4×10^2 J; (b) 2.4×10^2 J **107.** 15 J **109.** (a) 2.35×10^3 J; (b) 352 J **111.** 738 m **113.** (a) –3.8 kJ; (b) 31 kN **115.** (a) 300 J; (b) 93.8 J; (c) 6.38 m **117.** (a) 5.6 J; (b) 12 J; (c) 13 J **119.** (a) 1.2 J; (b) 11 m/s; (c) no; (d) no **121.** (a) 2.1×10^6 kg; (b) $(100 + 1.5t)^{0.5}$ m/s; (c) $(1.5 \times 10^6)/(100 + 1.5t)^{0.5}$ N; (d) 6.7 km

CHAPTER 9

- CP** **1.** (a) origin; (b) fourth quadrant; (c) on y axis below origin; (d) origin; (e) third quadrant; (f) origin **2.** (a)– (c) at the center of mass, still at the origin (their forces are internal to the system and cannot move the center of mass) **3.** (Consider slopes and Eq. 9-23.) (a) 1, 3, and then 2 and 4 tie (zero force); (b) 3 **4.** (a) unchanged; (b) unchanged (see Eq. 9-32); (c) decrease (Eq. 9-35) **5.** (a) zero; (b) positive (initial p_y down y ; final p_y up y); (c) positive direction of y **6.** (No net external force; \vec{P} conserved.) (a) 0; (b) no; (c) $-x$ **7.** (a) 10 kg · m/s; (b) 14 kg · m/s; (c) 6 kg · m/s **8.** (a) 4 kg · m/s; (b) 8 kg · m/s; (c) 3 J **9.** (a) 2 kg · m/s (conserve momentum along x); (b) 3 kg · m/s (conserve momentum along y)
Q **1.** (a) 2 N, rightward; (b) 2 N, rightward; (c) greater than 2 N, rightward **3.** b, c, a **5.** (a) x yes, y no; (b) x yes, y no; (c) x no, y yes **7.** (a) c , kinetic energy cannot be negative; d , total kinetic energy

cannot increase; (b) a ; (c) b **9.** (a) one was stationary; (b) 2; (c) 5; (d) equal (pool player's result)

- 11.** (a) C ; (b) B ; (c) 3
P **1.** (a) –1.50 m; (b) –1.43 m **3.** (a) –6.5 cm; (b) 8.3 cm; (c) 1.4 cm **5.** (a) –0.45 cm; (b) –2.0 cm **7.** (a) 0; (b) 3.13×10^{-11} m **9.** (a) 28 cm; (b) 2.3 m/s **11.** $(-4.0 \text{ m})\hat{i} + (4.0 \text{ m})\hat{j}$ **13.** 53 m **15.** (a) $(2.35\hat{i} - 1.57\hat{j}) \text{ m/s}^2$; (b) $(2.35\hat{i} - 1.57\hat{j})t \text{ m/s}$, with t in seconds; (d) straight, at downward angle 34° **17.** 4.2 m **19.** (a) 7.5×10^4 J; (b) 3.8×10^4 kg · m/s; (c) 39° south of due east **21.** (a) 5.0 kg · m/s; (b) 10 kg · m/s **23.** 1.0×10^3 to 1.2×10^3 kg · m/s **25.** (a) 42 N · s; (b) 2.1 kN **27.** (a) 67 m/s; (b) $-x$; (c) 1.2 kN; (d) $-x$ **29.** 5 N **31.** (a) 2.39×10^3 N · s; (b) 4.78×10^5 N; (c) 1.76×10^3 N · s; (d) 3.52×10^5 N **33.** (a) 5.86 kg · m/s; (b) 59.8°; (c) 2.93 kN; (d) 59.8° **35.** 9.9 × 10^2 N **37.** (a) 9.0 kg · m/s; (b) 3.0 kN; (c) 4.5 kN; (d) 20 m/s **39.** 3.0 mm/s **41.** (a) $(-0.15 \text{ m/s})\hat{i}$; (b) 0.18 m **43.** 55 cm **45.** (a) $(1.00\hat{i} - 0.167\hat{j}) \text{ km/s}$; (b) 3.23 MJ **47.** (a) 14 m/s; (b) 45° **49.** 3.1×10^2 m/s **51.** (a) 721 m/s; (b) 937 m/s **53.** (a) 33%; (b) 23%; (c) decreases **55.** (a) +2.0 m/s; (b) –1.3 J; (c) +40 J; (d) system got energy from some source, such as a small explosion **57.** (a) 4.4 m/s; (b) 0.80 **59.** 25 cm **61.** (a) 99 g; (b) 1.9 m/s; (c) 0.93 m/s **63.** (a) 3.00 m/s; (b) 6.00 m/s **65.** (a) 1.2 kg; (b) 2.5 m/s **67.** –28 cm **69.** (a) 0.21 kg; (b) 7.2 m **71.** (a) 4.15×10^5 m/s; (b) 4.84×10^5 m/s **73.** 120° **75.** (a) 433 m/s; (b) 250 m/s **77.** (a) 46 N; (b) none **79.** (a) 1.57×10^6 N; (b) 1.35×10^5 kg; (c) 2.08 km/s **81.** (a) 7290 m/s; (b) 8200 m/s; (c) 1.271×10^{10} J; (d) 1.275×10^{10} J **83.** (a) 1.92 m; (b) 0.640 m **85.** (a) 1.78 m/s; (b) less; (c) less; (d) greater **87.** (a) 3.7 m/s; (b) 1.3 N · s; (c) 1.8×10^2 N **89.** (a) $(7.4 \times 10^3 \text{ N} \cdot \text{s})\hat{i} - (7.4 \times 10^3 \text{ N} \cdot \text{s})\hat{j}$; (b) $(-7.4 \times 10^3 \text{ N} \cdot \text{s})\hat{i}$; (c) 2.3×10^3 N; (d) 2.1 × 10^4 N; (e) –45° **91.** +4.4 m/s **93.** 1.18×10^4 kg **95.** (a) 1.9 m/s; (b) –30°; (c) elastic **97.** (a) 6.9 m/s; (b) 30°; (c) 6.9 m/s; (d) –30°; (e) 2.0 m/s; (f) –180° **99.** (a) 25 mm; (b) 26 mm; (c) down; (d) 1.6×10^{-2} m/s² **101.** 29 J **103.** 2.2 kg **105.** 5.0 kg **107.** (a) 50 kg/s; (b) 1.6×10^2 kg/s **109.** (a) 4.6×10^3 km; (b) 73% **111.** 190 m/s **113.** 28.8 N **115.** (a) 0.745 mm; (b) 153°; (c) 1.67 mJ **117.** (a) $(2.67 \text{ m/s})\hat{i} + (-3.00 \text{ m/s})\hat{j}$; (b) 4.01 m/s; (c) 48.4° **119.** (a) –0.50 m; (b) –1.8 cm; (c) 0.50 m **121.** 0.22%

CHAPTER 10

- CP** **1.** b and c **2.** (a) and (d) ($\alpha = d^2\theta/dt^2$ must be a constant)
3. (a) yes; (b) no; (c) yes; (d) yes **4.** all tie **5.** 1, 2, 4, 3 (see Eq. 10-36) **6.** (see Eq. 10-40) 1 and 3 tie, 4, then 2 and 5 tie (zero) **7.** (a) downward in the figure ($\tau_{\text{net}} = 0$); (b) less (consider moment arms)
Q **1.** (a) c, a , then b and d tie; (b) b , then a and c tie, then d **3.** all tie **5.** (a) decrease; (b) clockwise; (c) counterclockwise **7.** larger **9.** c, a, b
P **1.** 14 rev **3.** (a) 4.0 rad/s; (b) 11.9 rad/s **5.** 11 rad/s **7.** (a) 4.0 m/s; (b) no **9.** (a) 3.00 s; (b) 18.9 rad **11.** (a) 30 s; (b) 1.8×10^3 rad **13.** (a) 3.4×10^2 s; (b) -4.5×10^{-3} rad/s²; (c) 98 s **15.** 8.0 s **17.** (a) 44 rad; (b) 5.5 s; (c) 32 s; (d) –2.1 s; (e) 40 s **19.** (a) 2.50×10^{-3} rad/s; (b) 20.2 m/s^2 ; (c) 0 **21.** 6.9×10^{-13} rad/s **23.** (a) 20.9 rad/s; (b) 12.5 m/s; (c) 800 rev/min²; (d) 600 rev **25.** (a) 7.3×10^{-5} rad/s; (b) 3.5×10^2 m/s; (c) 7.3×10^{-5} rad/s; (d) 4.6×10^2 m/s **27.** (a) 73 cm/s²; (b) 0.075; (c) 0.11 **29.** (a) 3.8×10^3 rad/s; (b) 1.9×10^2 m/s **31.** (a) 40 s; (b) 2.0 rad/s^2 **33.** 12.3 kg · m² **35.** (a) 1.1 kJ; (b) 9.7 kJ **37.** 0.097 kg · m² **39.** (a) 49 MJ; (b) 1.0×10^2 min **41.** (a) 0.023 kg · m²; (b) 1.1 mJ **43.** 4.7×10^{-4} kg · m² **45.** –3.85 N · m **47.** 4.6 N · m **49.** (a) 28.2 rad/s²; (b) 338 N · m **51.** (a) 6.00 cm/s²; (b) 4.87 N; (c) 4.54 N; (d) 1.20 rad/s²; (e) 0.0138 kg · m² **53.** 0.140 N **55.** 2.51×10^{-4} kg · m² **57.** (a) 4.2×10^2 rad/s²; (b) 5.0×10^2 rad/s **59.** 396 N · m **61.** (a) –19.8 kJ; (b) 1.32 kW **63.** 5.42 m/s **65.** (a) 5.32 m/s²; (b) 8.43 m/s²; (c) 41.8° **67.** 9.82 rad/s **69.** 6.16×10^{-5} kg · m² **71.** (a) 31.4 rad/s²; (b) 0.754 m/s²; (c) 56.1 N; (d) 55.1 N **73.** (a) 4.81×10^5 N; (b) 1.12×10^4 N · m; (c) 1.25×10^6 J **75.** (a) 2.3

AN-4 ANSWERS

rad/s^2 ; (b) 1.4 rad/s^2 **77.** (a) -67 rev/min^2 ; (b) 8.3 rev **81.** 3.1 rad/s **83.** (a) 1.57 m/s^2 ; (b) 4.55 N ; (c) 4.94 N **85.** 30 rev **87.** $0.054 \text{ kg} \cdot \text{m}^2$ **89.** $1.4 \times 10^2 \text{ N} \cdot \text{m}$ **91.** (a) 10 J ; (b) 0.27 m **93.** 4.6 rad/s^2 **95.** 2.6 J **97.** (a) $5.92 \times 10^4 \text{ m/s}^2$; (b) $4.39 \times 10^4 \text{ s}^{-2}$ **99.** (a) $0.791 \text{ kg} \cdot \text{m}^2$; (b) $1.79 \times 10^{-2} \text{ N} \cdot \text{m}$ **101.** (a) $1.5 \times 10^2 \text{ cm/s}$; (b) 15 rad/s ; (c) 15 rad/s ; (d) 75 cm/s ; (e) 3.0 rad/s **103.** (a) $7.0 \text{ kg} \cdot \text{m}^2$; (b) 7.2 m/s ; (c) 71°

CHAPTER 11

CP **1.** (a) same; (b) less **2.** less (consider the transfer of energy from rotational kinetic energy to gravitational potential energy) **3.** (draw the vectors, use right-hand rule) (a) $\pm z$; (b) $+y$; (c) $-x$ **4.** (see Eq. 11-21) (a) 1 and 3 tie; then 2 and 4 tie, then 5 (zero); (b) 2 and 3 **5.** (see Eqs. 11-23 and 11-16) (a) 3, 1; then 2 and 4 tie (zero); (b) 3 **6.** (a) all tie (same τ , same t , thus same ΔL); (b) sphere, disk, hoop (reverse order of I) **7.** (a) decreases; (b) same ($\tau_{\text{net}} = 0$, so L is conserved); (c) increases

Q **1.** *a*, then *b* and *c* tie, then *e*, *d* (zero) **3.** (a) spins in place; (b) rolls toward you; (c) rolls away from you **5.** (a) 1, 2, 3 (zero); (b) 1 and 2 tie, then 3; (c) 1 and 3 tie, then 2 **7.** (a) same; (b) increase; (c) decrease; (d) same, decrease, increase **9.** *D*, *B*, then *A* and *C* tie

P **1.** (a) 0; (b) $(22 \text{ m/s})\hat{i}$; (c) $(-22 \text{ m/s})\hat{i}$; (d) 0; (e) $1.5 \times 10^3 \text{ m/s}^2$; (f) $1.5 \times 10^3 \text{ m/s}^2$; (g) $(22 \text{ m/s})\hat{i}$; (h) $(44 \text{ m/s})\hat{i}$; (i) 0; (j) 0; (k) $1.5 \times 10^3 \text{ m/s}^2$; (l) $1.5 \times 10^3 \text{ m/s}^2$ **3.** -3.15 J **5.** 0.020 **7.** (a) 63 rad/s ; (b) 4.0 m **9.** 4.8 m **11.** (a) $(-4.0 \text{ N})\hat{i}$; (b) $0.60 \text{ kg} \cdot \text{m}^2$ **13.** 0.50 **15.** (a) $-(0.11 \text{ m})\omega$; (b) -2.1 m/s^2 ; (c) -47 rad/s^2 ; (d) 1.2 s ; (e) 8.6 m ; (f) 6.1 m/s **17.** (a) 13 cm/s^2 ; (b) 4.4 s ; (c) 55 cm/s ; (d) 18 mJ ; (e) 1.4 J ; (f) 27 rev/s **19.** $(-2.0 \text{ N} \cdot \text{m})\hat{i}$ **21.** (a) $(6.0 \text{ N} \cdot \text{m})\hat{j} + (8.0 \text{ N} \cdot \text{m})\hat{k}$; (b) $(-22 \text{ N} \cdot \text{m})\hat{i}$ **23.** (a) $(-1.5 \text{ N} \cdot \text{m})\hat{i} - (4.0 \text{ N} \cdot \text{m})\hat{j} - (1.0 \text{ N} \cdot \text{m})\hat{k}$; (b) $(-1.5 \text{ N} \cdot \text{m})\hat{i} - (4.0 \text{ N} \cdot \text{m})\hat{j} - (1.0 \text{ N} \cdot \text{m})\hat{k}$ **25.** (a) $(50 \text{ N} \cdot \text{m})\hat{k}$; (b) 90° **27.** (a) 0; (b) $(8.0 \text{ N} \cdot \text{m})\hat{i} + (8.0 \text{ N} \cdot \text{m})\hat{k}$ **29.** (a) $9.8 \text{ kg} \cdot \text{m}^2/\text{s}$; (b) $+z$ direction **31.** (a) 0; (b) $-22.6 \text{ kg} \cdot \text{m}^2/\text{s}$; (c) $-7.84 \text{ N} \cdot \text{m}$; (d) $-7.84 \text{ N} \cdot \text{m}$ **33.** (a) $(-1.7 \times 10^2 \text{ kg} \cdot \text{m}^2/\text{s})\hat{k}$; (b) $(+56 \text{ N} \cdot \text{m})\hat{k}$; (c) $(+56 \text{ kg} \cdot \text{m}^2/\text{s}^2)\hat{k}$ **35.** (a) $48\hat{k} \text{ N} \cdot \text{m}$; (b) increasing **37.** (a) $4.6 \times 10^{-3} \text{ kg} \cdot \text{m}^2$; (b) $1.1 \times 10^{-3} \text{ kg} \cdot \text{m}^2/\text{s}$; (c) $3.9 \times 10^{-3} \text{ kg} \cdot \text{m}^2/\text{s}$ **39.** (a) $1.47 \text{ N} \cdot \text{m}$; (b) 20.4 rad ; (c) -29.9 J ; (d) 19.9 W **41.** (a) $1.6 \text{ kg} \cdot \text{m}^2$; (b) $4.0 \text{ kg} \cdot \text{m}^2/\text{s}$ **43.** (a) 1.5 m ; (b) 0.93 rad/s ; (c) 98 J ; (d) 8.4 rad/s ; (e) $8.8 \times 10^2 \text{ J}$; (f) internal energy of the skaters **45.** (a) 3.6 rev/s ; (b) 3.0 ; (c) forces on the bricks from the man transferred energy from the man's internal energy to kinetic energy **47.** 0.17 rad/s **49.** (a) 750 rev/min ; (b) 450 rev/min ; (c) clockwise **51.** (a) 267 rev/min ; (b) 0.667 **53.** $1.3 \times 10^3 \text{ m/s}$ **55.** 3.4 rad/s **57.** (a) 18 rad/s ; (b) 0.92 **59.** 11.0 m/s **61.** 1.5 rad/s **63.** 0.070 rad/s **65.** (a) 0.148 rad/s ; (b) 0.0123 ; (c) 181° **67.** (a) 0.180 m ; (b) clockwise **69.** 0.041 rad/s **71.** (a) 1.6 m/s^2 ; (b) 16 rad/s^2 ; (c) $(4.0 \text{ N})\hat{i}$ **73.** (a) 0; (b) 0; (c) $-30t^3\hat{k} \text{ kg} \cdot \text{m}^2/\text{s}$; (d) $-90t^2\hat{k} \text{ N} \cdot \text{m}$; (e) $30t^3\hat{k} \text{ kg} \cdot \text{m}^2/\text{s}$; (f) $90t^2\hat{k} \text{ N} \cdot \text{m}$ **75.** (a) $149 \text{ kg} \cdot \text{m}^2$; (b) $158 \text{ kg} \cdot \text{m}^2/\text{s}$; (c) 0.744 rad/s **77.** (a) $6.65 \times 10^{-5} \text{ kg} \cdot \text{m}^2/\text{s}$; (b) no; (c) 0; (d) yes **79.** (a) 0.333 ; (b) 0.111 **81.** (a) 58.8 J ; (b) 39.2 J **83.** (a) 61.7 J ; (b) 3.43 m ; (c) no **85.** (a) $mvR/(I + MR^2)$; (b) $mvR^2/(I + MR^2)$ **87.** rotational speed would decrease; day would be about 0.8 s longer **89.** (a) 12.7 rad/s ; (b) clockwise **91.** (a) 0.81 mJ ; (b) 0.29 ; (c) $1.3 \times 10^{-2} \text{ N}$ **93.** (a) $mR^2/2$; (b) a solid circular cylinder

CHAPTER 12

CP **1.** *c, e, f* **2.** (a) no; (b) at site of \vec{F}_i , perpendicular to plane of figure; (c) 45 N **3.** *d*

Q **1.** (a) 1 and 3 tie, then 2; (b) all tie; (c) 1 and 3 tie, then 2 (zero) **3.** *a* and *c* (forces and torques balance) **5.** (a) 12 kg ; (b) 3 kg ; (c) 1 kg **7.** (a) at *C* (to eliminate forces there from a torque equation); (b) plus; (c) minus; (d) equal **9.** increase

P **1.** (a) 1.00 m ; (b) 2.00 m ; (c) 0.987 m ; (d) 1.97 m **3.** (a) 9.4 N ; (b) 4.4 N **5.** 7.92 kN **7.** (a) $2.8 \times 10^2 \text{ N}$; (b) $8.8 \times 10^2 \text{ N}$; (c) 71° **9.** 74.4 g **11.** (a) 1.2 kN ; (b) down; (c) 1.7 kN ; (d) up; (e) left; (f) right **13.** (a) 2.7 kN ; (b) up; (c) 3.6 kN ; (d) down **15.** (a) 5.0 N ; (b) 30 N ; (c) 1.3 m **17.** (a) 0.64 m ; (b) increased **19.** 8.7 N **21.** (a) 6.63 kN ; (b) 5.74 kN ; (c) 5.96 kN **23.** (a) 192 N ; (b) 96.1 N ; (c) 55.5 N **25.** 13.6 N **27.** (a) 1.9 kN ; (b) up; (c) 2.1 kN ; (d) down **29.** (a) $(-80 \text{ N})\hat{i} + (1.3 \times 10^2 \text{ N})\hat{j}$; (b) $(80 \text{ N})\hat{i} + (1.3 \times 10^2 \text{ N})\hat{j}$ **31.** 2.20 m **33.** (a) 60.0° ; (b) 300 N **35.** (a) 445 N ; (b) 0.50 ; (c) 315 N **37.** 0.34 **39.** (a) 211 N ; (b) 534 N ; (c) 320 N **41.** (a) slides; (b) 31° ; (c) tips; (d) 34° **43.** (a) $6.5 \times 10^6 \text{ N/m}^2$; (b) $1.1 \times 10^{-5} \text{ m}$ **45.** (a) 0.80 ; (b) 0.20 ; (c) 0.25 **47.** (a) $1.4 \times 10^9 \text{ N}$; (b) 75 **49.** (a) 866 N ; (b) 143 N ; (c) 0.165 **51.** (a) $1.2 \times 10^2 \text{ N}$; (b) 68 N **53.** (a) $1.8 \times 10^7 \text{ N}$; (b) $1.4 \times 10^7 \text{ N}$; (c) 16 **55.** 0.29 **57.** 76 N **59.** (a) 8.01 kN ; (b) 3.65 kN ; (c) 5.66 kN **61.** 71.7 N **63.** (a) $L/2$; (b) $L/4$; (c) $L/6$; (d) $L/8$; (e) $25L/24$ **65.** (a) 88 N ; (b) $(30\hat{i} + 97\hat{j}) \text{ N}$ **67.** $2.4 \times 10^9 \text{ N/m}^2$ **69.** 60° **71.** (a) $\mu < 0.57$; (b) $\mu > 0.57$ **73.** (a) $(35\hat{i} + 200\hat{j}) \text{ N}$; (b) $(-45\hat{i} + 200\hat{j}) \text{ N}$; (c) $1.9 \times 10^2 \text{ N}$ **75.** (a) BC, CD, DA ; (b) 535 N ; (c) 757 N **77.** (a) 1.38 kN ; (b) 180 N **79.** (a) $a_1 = L/2$, $a_2 = 5L/8$, $h = 9L/8$; (b) $b_1 = 2L/3$, $b_2 = L/2$, $h = 7L/6$ **81.** $L/4$ **83.** (a) 106 N ; (b) 64.0°

CHAPTER 13

CP **1.** all tie **2.** (a) 1, tie of 2 and 4, then 3; (b) line *d* **3.** (a) increase; (b) negative **4.** (a) 2; (b) 1 **5.** (a) path 1 (decreased E (more negative) gives decreased a); (b) less (decreased a gives decreased T)

Q **1.** $3GM^2/d^2$, leftward **3.** Gm^2/r^2 , upward **5.** *b* and *c* tie, then *a* (zero) **7.** 1, tie of 2 and 4, then 3 **9.** (a) positive y ; (b) yes, rotates counterclockwise until it points toward particle *B* **11.** *b*, *d*, and *f* all tie, then *e*, *c*, *a*

P **1.** $\frac{1}{2}$ **3.** 19 m **5.** 0.8 m **7.** $-5.00d$ **9.** $2.60 \times 10^5 \text{ km}$ **11.** (a) $M = m$; (b) 0 **13.** $8.31 \times 10^{-9} \text{ N}$ **15.** (a) $-1.88d$; (b) $-3.90d$; (c) $0.489d$ **17.** (a) 17 N ; (b) 2.4 **19.** $2.6 \times 10^6 \text{ m}$ **21.** $5 \times 10^{24} \text{ kg}$ **23.** (a) 7.6 m/s^2 ; (b) 4.2 m/s^2 **25.** (a) $(3.0 \times 10^{-7} \text{ N/kg})m$; (b) $(3.3 \times 10^{-7} \text{ N/kg})m$; (c) $(6.7 \times 10^{-7} \text{ N/kg})mr$ **27.** (a) 9.83 m/s^2 ; (b) 9.84 m/s^2 ; (c) 9.79 m/s^2 **29.** $5.0 \times 10^9 \text{ J}$ **31.** (a) 0.74 ; (b) 3.8 m/s^2 ; (c) 5.0 km/s **33.** (a) 0.0451 ; (b) 28.5 **35.** $-4.82 \times 10^{-13} \text{ J}$ **37.** (a) 0.50 pJ ; (b) -0.50 pJ **39.** (a) 1.7 km/s ; (b) $2.5 \times 10^5 \text{ m}$; (c) 1.4 km/s **41.** (a) 82 km/s ; (b) $1.8 \times 10^4 \text{ km/s}$ **43.** (a) 7.82 km/s ; (b) 87.5 min **45.** $6.5 \times 10^{23} \text{ kg}$ **47.** $5 \times 10^{10} \text{ stars}$ **49.** (a) $1.9 \times 10^{13} \text{ m}$; (b) $3.6R_p$ **51.** (a) $6.64 \times 10^3 \text{ km}$; (b) 0.0136 **53.** $5.8 \times 10^6 \text{ m}$ **57.** 0.71 y **59.** $(GM/L)^{0.5}$ **61.** (a) $3.19 \times 10^3 \text{ km}$; (b) lifting **63.** (a) 2.8 y ; (b) 1.0×10^{-4} **65.** (a) $r^{1.5}$; (b) r^{-1} ; (c) $r^{0.5}$; (d) $r^{-0.5}$ **67.** (a) 7.5 km/s ; (b) 97 min ; (c) $4.1 \times 10^2 \text{ km}$; (d) 7.7 km/s ; (e) 93 min ; (f) $3.2 \times 10^{-3} \text{ N}$; (g) no; (h) yes **69.** 1.1 s **71.** (a) $GMmx(x^2 + R^2)^{-3/2}$; (b) $[2GM(R^{-1} - (R^2 + x^2)^{-1/2})]^{1/2}$ **73.** (a) $1.0 \times 10^3 \text{ kg}$; (b) 1.5 km/s **75.** $3.2 \times 10^{-7} \text{ N}$ **77.** $0.037\hat{j} \mu\text{N}$ **79.** $2\pi r^{1.5}G^{-0.5}(M + m/4)^{-0.5}$ **81.** (a) $2.2 \times 10^{-7} \text{ rad/s}$; (b) 89 km/s **83.** (a) $2.15 \times 10^4 \text{ s}$; (b) 12.3 km/s ; (c) 12.0 km/s ; (d) $2.17 \times 10^{11} \text{ J}$; (e) $-4.53 \times 10^{11} \text{ J}$; (f) $-2.35 \times 10^{11} \text{ J}$; (g) $4.04 \times 10^7 \text{ m}$; (h) $1.22 \times 10^3 \text{ s}$; (i) elliptical **85.** $2.5 \times 10^4 \text{ km}$ **87.** (a) $1.4 \times 10^6 \text{ m/s}$; (b) $3 \times 10^6 \text{ m/s}^2$ **89.** (a) 0; (b) $1.8 \times 10^{32} \text{ J}$; (c) $1.8 \times 10^{32} \text{ J}$; (d) 0.99 km/s **91.** (a) Gm^2/R_i ; (b) $Gm^2/2R_i$; (c) $(Gm/R_i)^{0.5}$; (d) $2(Gm/R_i)^{0.5}$; (e) Gm^2/R_i ; (f) $(2Gm/R_i)^{0.5}$; (g) The center-of-mass frame is an inertial frame, and in it the principle of conservation of energy may be written as in Chapter 8; the reference frame attached to body *A* is noninertial, and the principle cannot be written as in Chapter 8. Answer (d) is correct. **93.** $2.4 \times 10^4 \text{ m/s}$ **95.** $-0.044\hat{j} \mu\text{N}$ **97.** $GM_E m/12R_E$

CHAPTER 14

- CP** 1. all tie 2. (a) all tie (the gravitational force on the penguin is the same); (b) $0.95\rho_0, \rho_0, 1.1\rho_0$ 3. $13 \text{ cm}^3/\text{s}$, outward
 4. (a) all tie; (b) 1, then 2 and 3 tie, 4 (wider means slower);
 (c) 4, 3, 2, 1 (wider and lower mean more pressure)
Q 1. (a) moves downward; (b) moves downward 3. (a) downward;
 (b) downward; (c) same 5. b , then a and d tie (zero), then c
 7. (a) 1 and 4; (b) 2; (c) 3 9. *B, C, A*
P 1. 0.074 3. $1.1 \times 10^5 \text{ Pa}$ 5. $2.9 \times 10^4 \text{ N}$ 7. (b) 26 kN
 9. (a) $1.0 \times 10^3 \text{ torr}$; (b) $1.7 \times 10^3 \text{ torr}$ 11. (a) 94 torr; (b) 4.1×10^2 torr; (c) $3.1 \times 10^2 \text{ torr}$ 13. $1.08 \times 10^3 \text{ atm}$ 15. $-2.6 \times 10^4 \text{ Pa}$
 17. $7.2 \times 10^5 \text{ N}$ 19. $4.69 \times 10^5 \text{ N}$ 21. 0.635 J 23. 44 km
 25. 739.26 torr 27. (a) 7.9 km; (b) 16 km 29. 8.50 kg 31. (a)
 $6.7 \times 10^2 \text{ kg/m}^3$; (b) $7.4 \times 10^2 \text{ kg/m}^3$ 33. (a) $2.04 \times 10^{-2} \text{ m}^3$;
 (b) 1.57 kN 35. five 37. 57.3 cm 39. (a) 1.2 kg ; (b) 1.3×10^3 kg/m^3 41. (a) 0.10; (b) 0.083 43. (a) 637.8 cm^3 ; (b) 5.102 m^3 ;
 (c) $5.102 \times 10^3 \text{ kg}$ 45. 0.126 m^3 47. (a) 1.80 m^3 ; (b) 4.75 m^3
 49. (a) 3.0 m/s ; (b) 2.8 m/s 51. 8.1 m/s 53. 66 W 55. $1.4 \times 10^5 \text{ J}$
 57. (a) $1.6 \times 10^{-3} \text{ m}^3/\text{s}$; (b) 0.90 m 59. (a) 2.5 m/s ; (b) $2.6 \times 10^5 \text{ Pa}$
 61. (a) 3.9 m/s ; (b) 88 kPa 63. $1.1 \times 10^2 \text{ m/s}$ 65. (b) 2.0×10^{-2} m^3/s 67. (a) 74 N ; (b) $1.5 \times 10^2 \text{ m}^3$ 69. (a) $0.0776 \text{ m}^3/\text{s}$; (b) 69.8 kg/s 71. (a) 35 cm; (b) 30 cm; (c) 20 cm 73. 1.5 g/cm^3 75. $5.11 \times 10^{-7} \text{ kg}$ 77. 44.2 g 79. $6.0 \times 10^2 \text{ kg/m}^3$ 81. 45.3 cm^3
 83. (a) 3.2 m/s ; (b) $9.2 \times 10^4 \text{ Pa}$; (c) 10.3 m 85. $1.07 \times 10^3 \text{ g}$

CHAPTER 15

- CP** 1. (sketch x versus t) (a) $-x_m$; (b) $+x_m$; (c) 0 2. a (F must have the form of Eq. 15-10) 3. (a) 5 J ; (b) 2 J ; (c) 5 J 4. all tie (in Eq. 15-29, m is included in I) 5. 1, 2, 3 (the ratio m/b matters; k does not)
Q 1. a and b 3. (a) 2; (b) positive; (c) between 0 and $+x_m$
 5. (a) between D and E ; (b) between $3\pi/2$ rad and 2π rad
 7. (a) all tie; (b) 3, then 1 and 2 tie; (c) 1, 2, 3 (zero); (d) 1, 2, 3 (zero);
 (e) 1, 3, 2 9. b (infinite period, does not oscillate), c, a
 11. (a) greater; (b) same; (c) same; (d) greater; (e) greater
P 1. (a) 0.50 s ; (b) 2.0 Hz ; (c) 18 cm 3. 37.8 m/s^2 5. (a) 1.0 mm ;
 (b) 0.75 m/s ; (c) $5.7 \times 10^2 \text{ m/s}^2$ 7. (a) 498 Hz ; (b) greater
 9. (a) 3.0 m ; (b) -49 m/s ; (c) $-2.7 \times 10^2 \text{ m/s}^2$; (d) 20 rad ; (e) 1.5 Hz ;
 (f) 0.67 s 11. 39.6 Hz 13. (a) 0.500 s ; (b) 2.00 Hz ; (c) 12.6 rad/s ;
 (d) 79.0 N/m ; (e) 4.40 m/s ; (f) 27.6 N 15. (a) $0.18A$; (b) same
 direction 17. (a) 5.58 Hz ; (b) 0.325 kg ; (c) 0.400 m 19. (a) 25 cm ;
 (b) 2.2 Hz 21. 54 Hz 23. 3.1 cm 25. (a) 0.525 m ; (b) 0.686 s
 27. (a) 0.75 ; (b) 0.25 ; (c) $2^{-0.5}x_m$ 29. 37 mJ 31. (a) 2.25 Hz ;
 (b) 125 J ; (c) 250 J ; (d) 86.6 cm 33. (a) 1.1 m/s ; (b) 3.3 cm
 35. (a) 3.1 ms ; (b) 4.0 m/s ; (c) 0.080 J ; (d) 80 N ; (e) 40 N
 37. (a) 2.2 Hz ; (b) 56 cm/s ; (c) 0.10 kg ; (d) 20.0 cm 39. (a) 39.5 rad/s ;
 (b) 34.2 rad/s ; (c) 124 rad/s^2 41. (a) $0.205 \text{ kg}\cdot\text{m}^2$; (b) 47.7 cm ;
 (c) 1.50 s 43. (a) 1.64 s ; (b) equal 45. 8.77 s 47. 0.366 s
 49. (a) 0.845 rad ; (b) 0.0602 rad 51. (a) 0.53 m ; (b) 2.1 s
 53. 0.0653 s 55. (a) 2.26 s ; (b) increases; (c) same 57. 6.0%
 59. (a) 14.3 s ; (b) 5.27 61. (a) $F_m/b\omega$; (b) F_m/b 63. 5.0 cm
 65. (a) $2.8 \times 10^3 \text{ rad/s}$; (b) 2.1 m/s ; (c) 5.7 km/s^2 67. (a) 1.1 Hz ;
 (b) 5.0 cm 69. 7.2 m/s 71. (a) 7.90 N/m ; (b) 1.19 cm ; (c) 2.00 Hz
 73. (a) $1.3 \times 10^2 \text{ N/m}$; (b) 0.62 s ; (c) 1.6 Hz ; (d) 5.0 cm ; (e) 0.51 m/s
 75. (a) 16.6 cm ; (b) 1.23% 77. (a) 1.2 J ; (b) 50 79. 1.53 m
 81. (a) 0.30 m ; (b) 0.28 s ; (c) $1.5 \times 10^2 \text{ m/s}^2$; (d) 11 J 83. (a) 1.23 kN/m ; (b) 76.0 N 85. 1.6 kg 87. (a) $0.735 \text{ kg}\cdot\text{m}^2$; (b) $0.0240 \text{ N}\cdot\text{m}$;
 (c) 0.181 rad/s 89. (a) 3.5 m ; (b) 0.75 s 91. (a) 0.35 Hz ; (b) 0.39 Hz ;
 (c) 0 (no oscillation) 93. (a) 245 N/m ; (b) 0.284 s

95. $0.079 \text{ kg}\cdot\text{m}^2$ 97. (a) $8.11 \times 10^{-5} \text{ kg}\cdot\text{m}^2$; (b) 3.14 rad/s

99. 14.0° 101. (a) 3.2 Hz ; (b) 0.26 m ; (c) $x = (0.26 \text{ m}) \cos(20t - \pi/2)$, with t in seconds 103. (a) 0.44 s ; (b) 0.18 m
 105. (a) 0.45 s ; (b) 0.10 m above and 0.20 m below; (c) 0.15 m ;
 (d) 2.3 J 107. $7 \times 10^2 \text{ N/m}$

CHAPTER 16

- CP** 1. a, 2; b, 3; c, 1 (compare with the phase in Eq. 16-2, then see Eq. 16-5) 2. (a) 2, 3, 1 (see Eq. 16-12); (b) 3, then 1 and 2 tie (find amplitude of dy/dt) 3. (a) same (independent of f); (b) decrease ($\lambda = v/f$); (c) increase; (d) increase 4. 0.20 and 0.80 tie, then 0.60, 0.45 5. (a) 1; (b) 3; (c) 2 6. (a) 75 Hz ; (b) 525 Hz

Q 1. (a) 1, 4, 2, 3; (b) 1, 4, 2, 3 3. a , upward; b , upward; c , downward; d , downward; e , downward; f , downward; g , upward; h , upward

5. intermediate (closer to fully destructive) 7. (a) 0, 0.2 wavelength, 0.5 wavelength (zero); (b) $4P_{\text{avg},1}$ 9. d 11. c, a, b

- P** 1. 1.1 ms 3. (a) 3.49 m^{-1} ; (b) 31.5 m/s 5. (a) 0.680 s ; (b) 1.47 Hz ; (c) 2.06 m/s 7. (a) 64 Hz ; (b) 1.3 m ; (c) 4.0 cm ; (d) 5.0 m^{-1} ;
 (e) $4.0 \times 10^2 \text{ s}^{-1}$; (f) $\pi/2 \text{ rad}$; (g) minus 9. (a) 3.0 mm ; (b) 16 m^{-1} ;
 (c) $2.4 \times 10^2 \text{ s}^{-1}$; (d) minus 11. (a) negative; (b) 4.0 cm ; (c) 0.31 cm^{-1} ; (d) 0.63 s^{-1} ; (e) $\pi \text{ rad}$; (f) minus; (g) 2.0 cm/s ; (h) -2.5 cm/s 13. (a) 11.7 cm ; (b) $\pi \text{ rad}$ 15. (a) 0.12 mm ; (b) 141 m^{-1} ; (c) 628 s^{-1} ;
 (d) plus 17. (a) 15 m/s ; (b) 0.036 N 19. 129 m/s 21. 2.63 m
 23. (a) 5.0 cm ; (b) 40 cm ; (c) 12 m/s ; (d) 0.033 s ; (e) 9.4 m/s ;
 (f) 16 m^{-1} ; (g) $1.9 \times 10^2 \text{ s}^{-1}$; (h) 0.93 rad ; (i) plus 27. 3.2 mm
 29. 0.20 m/s 31. $1.41y_m$ 33. (a) 9.0 mm ; (b) 16 m^{-1} ; (c) $1.1 \times 10^3 \text{ s}^{-1}$;
 (d) 2.7 rad ; (e) plus 35. 5.0 cm 37. (a) 3.29 mm ; (b) 1.55 rad ;
 (c) 1.55 rad 39. 84° 41. (a) 82.0 m/s ; (b) 16.8 m ; (c) 4.88 Hz
 43. (a) 7.91 Hz ; (b) 15.8 Hz ; (c) 23.7 Hz 45. (a) 105 Hz ; (b) 158 m/s
 47. 260 Hz 49. (a) 144 m/s ; (b) 60.0 cm ; (c) 241 Hz 51. (a) 0.50 cm ;
 (b) 3.1 m^{-1} ; (c) $3.1 \times 10^2 \text{ s}^{-1}$; (d) minus 53. (a) 0.25 cm ; (b) $1.2 \times 10^2 \text{ cm/s}$; (c) 3.0 cm ; (d) 0 55. 0.25 m 57. (a) 2.00 Hz ; (b) 2.00 m ; (c) 4.00 m/s ; (d) 50.0 cm ; (e) 150 cm ; (f) 250 cm ; (g) 0; (h) 100 cm ; (i) 200 cm
 59. (a) 324 Hz ; (b) eight 61. 36 N 63. (a) 75 Hz ; (b) 13 ms
 65. (a) 2.0 mm ; (b) 95 Hz ; (c) $+30 \text{ m/s}$; (d) 31 cm ; (e) 1.2 m/s
 67. (a) 0.31 m ; (b) 1.64 rad ; (c) 2.2 mm 69. (a) $0.83y_1$; (b) 37°
 71. (a) 3.77 m/s ; (b) 12.3 N ; (c) 0; (d) 46.4 W ; (e) 0; (f) 0; (g) $\pm 0.50 \text{ cm}$
 73. 1.2 rad 75. (a) 300 m/s ; (b) no 77. (a) $[k \Delta\ell(\ell + \Delta\ell)/m]^{0.5}$
 79. (a) 144 m/s ; (b) 3.00 m ; (c) 1.50 m ; (d) 48.0 Hz ; (e) 96.0 Hz
 81. (a) 1.00 cm ; (b) $3.46 \times 10^3 \text{ s}^{-1}$; (c) 10.5 m^{-1} ; (d) plus 83. (a)
 $2\pi y_m/\lambda$; (b) no 85. (a) 240 cm ; (b) 120 cm ; (c) 80 cm 87. (a) 1.33 m/s ; (b) 1.88 m/s ; (c) 16.7 m/s^2 ; (d) 23.7 m/s^2 89. (a) 0.52 m ; (b) 40 m/s ; (c) 0.40 m 91. (a) 0.16 m ; (b) $2.4 \times 10^2 \text{ N}$; (c) $y(x, t) = (0.16 \text{ m}) \sin[(1.57 \text{ m}^{-1})x] \sin[(31.4 \text{ s}^{-1})t]$ 93. (c) 2.0 m/s ; (d) $-x$

CHAPTER 17

- CP** 1. beginning to decrease (example: mentally move the curves of Fig. 17-6 rightward past the point at $x = 42 \text{ cm}$) 2. (a) 1 and 2 tie, then 3 (see Eq. 17-28); (b) 3, then 1 and 2 tie (see Eq. 17-26)
 3. second (see Eqs. 17-39 and 17-41) 4. a , greater; b , less; c , can't tell; d , can't tell; e , greater; f , less

- Q** 1. (a) 0, 0.2 wavelength, 0.5 wavelength (zero); (b) $4P_{\text{avg},1}$
 3. C , then A and B tie 5. E, A, D, C, B 7. 1, 4, 3, 2 9. 150 Hz and 450 Hz
P 1. (a) 79 m ; (b) 41 m ; (c) 89 m 3. (a) 2.6 km ; (b) $2.0 \times 10^2 \text{ m}$ 5. $1.9 \times 10^3 \text{ km}$ 7. 40.7 m 9. 0.23 ms 11. (a) $76.2 \mu\text{m}$; (b) 0.333 mm
 13. 960 Hz 15. (a) $2.3 \times 10^2 \text{ Hz}$; (b) higher 17. (a) 143 Hz ; (b) 3;
 (c) 5; (d) 286 Hz ; (e) 2; (f) 3 19. (a) 14; (b) 14 21. (a) 343 Hz ;
 (b) 3; (c) 5; (d) 686 Hz ; (e) 2; (f) 3 23. (a) 0; (b) fully constructive;
 (c) increase; (d) 128 m ; (e) 63.0 m ; (f) 41.2 m 25. 36.8 nm

AN-6 ANSWERS

- 27.** (a) 1.0×10^3 ; (b) 32 **29.** 15.0 mW **31.** 2 μW **33.** 0.76 μm
35. (a) 5.97×10^{-5} W/m²; (b) 4.48 nW **37.** (a) 0.34 nW; (b) 0.68 nW;
(c) 1.4 nW; (d) 0.88 nW; (e) 0 **39.** (a) 405 m/s; (b) 596 N; (c) 44.0
cm; (d) 37.3 cm **41.** (a) 833 Hz; (b) 0.418 m **43.** (a) 3; (b) 1129 Hz;
(c) 1506 Hz **45.** (a) 2; (b) 1 **47.** 12.4 m **49.** 45.3 N **51.** 2.25 ms
53. 0.020 **55.** (a) 526 Hz; (b) 555 Hz **57.** 0 **59.** (a) 1.022 kHz;
(b) 1.045 kHz **61.** 41 kHz **63.** 155 Hz **65.** (a) 2.0 kHz; (b) 2.0
kHz **67.** (a) 485.8 Hz; (b) 500.0 Hz; (c) 486.2 Hz; (d) 500.0 Hz
69. (a) 42°; (b) 11 s **71.** 1 cm **73.** 2.1 m **75.** (a) 39.7 $\mu\text{W}/\text{m}^2$;
(b) 171 nm; (c) 0.893 Pa **77.** 0.25 **79.** (a) 2.10 m; (b) 1.47 m
81. (a) 59.7; (b) 2.81×10^{-4} **83.** (a) rightward; (b) 0.90 m/s; (c) less
85. (a) 11 ms; (b) 3.8 m **87.** (a) 9.7×10^2 Hz; (b) 1.0 kHz; (c) 60 Hz,
no **89.** (a) 21 nm; (b) 35 cm; (c) 24 nm; (d) 35 cm **91.** (a) 7.70 Hz;
(b) 7.70 Hz **93.** (a) 5.2 kHz; (b) 2 **95.** (a) 10 W; (b) 0.032 W/m²;
(c) 99 dB **97.** (a) 0; (b) 0.572 m; (c) 1.14 m **99.** 171 m **101.** (a)
 3.6×10^2 m/s; (b) 150 Hz **103.** 400 Hz **105.** (a) 14; (b) 12

CHAPTER 18

- CP** 1. (a) all tie; (b) 50°X, 50°Y, 50°W **2.** (a) 2 and 3 tie, then 1,
then 4; (b) 3, 2, then 1 and 4 tie (from Eqs. 18-9 and 18-10, assume
that change in area is proportional to initial area) **3.** A (see
Eq. 18-14) **4.** c and e (maximize area enclosed by a clockwise
cycle) **5.** (a) all tie (ΔE_{int} depends on *i* and *f*, not on path); (b) 4, 3,
2, 1 (compare areas under curves); (c) 4, 3, 2, 1 (see Eq. 18-26)
6. (a) zero (closed cycle); (b) negative (W_{net} is negative; see
Eq. 18-26) **7.** b and d tie, then a, c (P_{cond} identical; see Eq. 18-32)

- Q** 1. c, then the rest tie **3.B**, then A and C tie **5.** (a) *f*, because
ice temperature will not rise to freezing point and then drop; (b) *b*
and *c* at freezing point, *d* above, *e* below; (c) in *b* liquid partly
freezes and no ice melts; in *c* no liquid freezes and no ice melts; in *d*
no liquid freezes and ice fully melts; in *e* liquid fully freezes and no
ice melts **7.** (a) both clockwise; (b) both clockwise **9.** (a) greater;
(b) 1, 2, 3; (c) 1, 3, 2; (d) 1, 2, 3; (e) 2, 3, 1 **11.** c, b, a

- P** 1. 1.366 **3.** 348 K **5.** (a) 320°F; (b) -12.3°F **7.** -92.1°X
9.2731 cm **11.** 49.87 cm³ **13.** 29 cm³ **15.** 360°C **17.** 0.26 cm³
19. 0.13 mm **21.** 7.5 cm **23.** 160 s **25.** 94.6 L **27.** 42.7 kJ
29. 33 m² **31.** 33 g **33.** 3.0 min **35.** 13.5 C° **37.** (a) 5.3°C; (b) 0;
(c) 0°C; (d) 60 g **39.** 742 kJ **41.** (a) 0°C; (b) 2.5°C **43.** (a) $1.2 \times$
 10^2 J; (b) 75 J; (c) 30 J **45.** -30 J **47.** (a) 6.0 cal; (b) -43 cal;
(c) 40 cal; (d) 18 cal; (e) 18 cal **49.** 60 J **51.** (a) 1.23 kW; (b)
2.28 kW; (c) 1.05 kW **53.** 1.66 kJ/s **55.** (a) 16 J/s; (b) 0.048 g/s
57. (a) 1.7×10^4 W/m²; (b) 18 W/m² **59.** 0.50 min **61.** 0.40 cm/h
63. -4.2°C **65.** 1.1 m **67.** 10% **69.** (a) 80 J; (b) 80 J **71.** 4.5 \times
 10^2 J/kg·K **73.** 0.432 cm³ **75.** 3.1×10^2 J **77.** 79.5°C **79.** 23 J
81. (a) $11p_1V_1$; (b) $6p_1V_1$ **83.** 4.83×10^{-2} cm³ **85.** 10.5°C
87. (a) 90 W; (b) 2.3×10^2 W; (c) 3.3×10^2 W **89.** (a) 1.87×10^4 ;
(b) 10.4 h **91.** 333 J **93.** 8.6 J **95.** (a) -45 J; (b) +45 J

CHAPTER 19

- CP** 1. all but *c* **2.** (a) all tie; (b) 3, 2, 1 **3.** gas *A* **4.** 5 (greatest
change in *T*), then tie of 1, 2, 3, and 4 **5.** 1, 2, 3 ($Q_3 = 0$, Q_2 goes
into work W_2 , but Q_1 goes into greater work W_1 and increases
gas temperature)

- Q** 1. *d*, then *a* and *b* tie, then *c* **3.** 20 J **5.** (a) 3; (b) 1; (c) 4; (d) 2;
(e) yes **7.** (a) 1, 2, 3, 4; (b) 1, 2, 3 **9.** constant-volume process

- P** 1. 0.933 kg **3.** (a) 0.0388 mol; (b) 220°C **5.** 25 molecules/cm³
7. (a) 3.14×10^3 J; (b) from **9.** 186 kPa **11.** 5.60 kJ
13. (a) 1.5 mol; (b) 1.8×10^3 K; (c) 6.0×10^2 K; (d) 5.0 kJ

- 15.** 360 K **17.** 2.0×10^5 Pa **19.** (a) 511 m/s; (b) -200°C; (c) 899°C
21. 1.8×10^2 m/s **23.** 1.9 kPa **25.** (a) 5.65×10^{-21} J; (b) 7.72 \times
 10^{-21} J; (c) 3.40 kJ; (d) 4.65 kJ **27.** (a) 6.76×10^{-20} J; (b) 10.7
km **29.** (a) 6×10^9 km **31.** (a) 3.27×10^{10} molecules/cm³; (b) 172 m
33. (a) 6.5 km/s; (b) 7.1 km/s **35.** (a) 420 m/s; (b) 458 m/s; (c) yes
37. (a) 0.67; (b) 1.2; (c) 1.3; (d) 0.33 **39.** (a) 1.0×10^4 K; (b) 1.6 \times
 10^5 K; (c) 4.4×10^2 K; (d) 7.0×10^3 K; (e) no; (f) yes **41.** (a) 7.0
km/s; (b) 2.0×10^{-8} cm; (c) 3.5×10^{10} collisions/s **43.** (a) 3.49 kJ;
(b) 2.49 kJ; (c) 997 J; (d) 1.00 kJ **45.** (a) 6.6×10^{-26} kg; (b) 40
g/mol **47.** (a) 0; (b) +374 J; (c) +374 J; (d) $+3.11 \times 10^{-22}$ J
49. 15.8 J/mol·K **51.** 8.0 kJ **53.** (a) 6.98 kJ; (b) 4.99 kJ; (c) 1.99 kJ;
(d) 2.99 kJ **55.** (a) 14 atm; (b) 6.2×10^2 K **57.** (a) diatomic;
(b) 446 K; (c) 8.10 mol **59.** -15 J **61.** -20 J **63.** (a) 3.74 kJ;
(b) 3.74 kJ; (c) 0; (d) 0; (e) -1.81 kJ; (f) 1.81 kJ; (g) -3.22 kJ;
(h) -1.93 kJ; (i) -1.29 kJ; (j) 520 J; (k) 0; (l) 520 J; (m) 0.0246 m³;
(n) 2.00 atm; (o) 0.0373 m³; (p) 1.00 atm **65.** (a) monatomic;
(b) 2.7×10^4 K; (c) 4.5×10^4 mol; (d) 3.4 kJ; (e) 3.4×10^2 kJ;
(f) 0.010 **67.** (a) 2.00 atm; (b) 333 J; (c) 0.961 atm; (d) 236 J
69. 349 K **71.** (a) -374 J; (b) 0; (c) +374 J; (d) $+3.11 \times 10^{-22}$ J
73. 7.03×10^9 s⁻¹ **75.** (a) 900 cal; (b) 0; (c) 900 cal; (d) 450 cal;
(e) 1200 cal; (f) 300 cal; (g) 900 cal; (h) 450 cal; (i) 0; (j) -900 cal;
(k) 900 cal; (l) 450 cal **77.** (a) $3/v_0^3$; (b) $0.750v_0$; (c) $0.775v_0$
79. (a) -2.37 kJ; (b) 2.37 kJ **81.** (b) 125 J; (c) to **83.** (a) 8.0 atm;
(b) 300 K; (c) 4.4 kJ; (d) 3.2 atm; (e) 120 K; (f) 2.9 kJ; (g) 4.6 atm;
(h) 170 K; (i) 3.4 kJ **85.** (a) 38 L; (b) 71 g **87.** -3.0 J

CHAPTER 20

- CP** 1. a, b, c **2.** smaller (*Q* is smaller) **3.** c, b, a **4.** a, d, c, b **5.** b

- Q** 1. b, a, c, d **3.** unchanged **5.** *a* and *c* tie, then *b* and *d* tie

- 7.** (a) same; (b) increase; (c) decrease **9.** A, first; B, first and
second; C, second; D, neither

- P** 1. (a) 9.22 kJ; (b) 23.1 J/K; (c) 0 **3.** 14.4 J/K **5.** (a) $5.79 \times$
 10^4 J; (b) 173 J/K **7.** (a) 320 K; (b) 0; (c) +1.72 J/K **9.** +0.76 J/K
11. (a) 57.0°C; (b) -22.1 J/K; (c) +24.9 J/K; (d) +2.8 J/K
13. (a) -710 mJ/K; (b) +710 mJ/K; (c) +723 mJ/K; (d) -723 mJ/K;
(e) +13 mJ/K; (f) 0 **15.** (a) -943 J/K; (b) +943 J/K; (c) yes
17. (a) 0.333; (b) 0.215; (c) 0.644; (d) 1.10; (e) 1.10; (f) 0; (g) 1.10;
(h) 0; (i) -0.889; (j) -0.889; (k) -1.10; (l) -0.889; (m) 0; (n) 0.889;
(o) 0 **19.** (a) 0.693; (b) 4.50; (c) 0.693; (d) 0; (e) 4.50; (f) 23.0 J/K;
(g) -0.693; (h) 7.50; (i) -0.693; (j) 3.00; (k) 4.50; (l) 23.0 J/K
21. -1.18 J/K **23.** 97 K **25.** (a) 266 K; (b) 341 K **27.** (a) 23.6%;
(b) 1.49×10^4 J **29.** (a) 2.27 kJ; (b) 14.8 kJ; (c) 15.4%; (d) 75.0%;
(e) greater **31.** (a) 33 kJ; (b) 25 kJ; (c) 26 kJ; (d) 18 kJ
33. (a) 1.47 kJ; (b) 554 J; (c) 918 J; (d) 62.4% **35.** (a) 3.00; (b) 1.98;
(c) 0.660; (d) 0.495; (e) 0.165; (f) 34.0% **37.** 440 W **39.** 20 J
41. 0.25 hp **43.** 2.03 **47.** (a) $W = N!/(n_1! n_2! n_3!)$; (b)
 $[(N/2)! (N/2)!]/[(N/3)! (N/3)!]$; (c) 4.2×10^{16} **49.** 0.141 J/K·s
51. (a) 87 m/s; (b) 1.2×10^2 m/s; (c) 22 J/K **53.** (a) 78%; (b) 82 kg/s
55. (a) 40.9°C; (b) -27.1 J/K; (c) 30.3 J/K; (d) 3.18 J/K **57.** +3.59 J/K
59. 1.18×10^3 J/K **63.** (a) 0; (b) 0; (c) -23.0 J/K; (d) 23.0 J/K
65. (a) 25.5 kJ; (b) 4.73 kJ; (c) 18.5% **67.** (a) 1.95 J/K; (b) 0.650 J/K;
(c) 0.217 J/K; (d) 0.072 J/K; (e) decrease **69.** (a) 4.45 J/K; (b) no
71. (a) 1.26×10^{14} ; (b) 4.71×10^{13} ; (c) 0.37; (d) 1.01×10^{29} ;
(e) 1.37×10^{28} ; (f) 0.14; (g) 9.05×10^{58} ; (h) 1.64×10^{57} ; (i) 0.018;
(j) decrease **73.** (a) 42.6 kJ; (b) 7.61 kJ **75.** (a) 1; (b) 1; (c) 3;
(d) 10; (e) 1.5×10^{-23} J/K; (f) 3.2×10^{-23} J/K

CHAPTER 21

- CP** 1. C and D attract; B and D attract 2. (a) leftward;
(b) leftward; (c) leftward 3. (a) a, c, b ; (b) less than $4. -15e$
(net charge of $-30e$ is equally shared)
- Q** 1. 3, 1, 2, 4 (zero) 3. a and b 5. $2kq^2/r^2$, up the page
7. b and c tie, then a (zero) 9. (a) same; (b) less than; (c) cancel;
(d) add; (e) adding components; (f) positive direction of y ;
(g) negative direction of y ; (h) positive direction of x ; (i) negative
direction of x
- P** 1. 0.500 3. 1.39 m 5. 2.81 N 7. -4.00 9. (a) $-1.00\ \mu\text{C}$;
(b) $3.00\ \mu\text{C}$ 11. (a) $0.17\ \text{N}$; (b) $-0.046\ \text{N}$ 13. (a) $-14\ \text{cm}$; (b) 0
15. (a) $35\ \text{N}$; (b) -10° ; (c) $-8.4\ \text{cm}$; (d) $+2.7\ \text{cm}$ 17. (a) $1.60\ \text{N}$;
(b) $2.77\ \text{N}$ 19. (a) $3.00\ \text{cm}$; (b) 0; (c) -0.444 21. $3.8 \times 10^{-8}\ \text{C}$
23. (a) 0; (b) $12\ \text{cm}$; (c) 0; (d) $4.9 \times 10^{-26}\ \text{N}$ 25. 6.3×10^{11}
27. (a) $3.2 \times 10^{-19}\ \text{C}$; (b) 2 29. (a) $-6.05\ \text{cm}$; (b) $6.05\ \text{cm}$
31. $122\ \text{mA}$ 33. $1.3 \times 10^7\ \text{C}$ 35. (a) 0; (b) $1.9 \times 10^{-9}\ \text{N}$
37. (a) 9B ; (b) ^{13}N ; (c) ^{12}C 39. $1.31 \times 10^{-22}\ \text{N}$ 41. (a) $5.7 \times 10^{13}\ \text{C}$;
(b) cancels out; (c) $6.0 \times 10^5\ \text{kg}$ 43. (b) $3.1\ \text{cm}$ 45. $0.19\ \text{MC}$
47. $-45\ \mu\text{C}$ 49. $3.8\ \text{N}$ 51. (a) 2.00×10^{10} electrons; (b) 1.33×10^{10}
electrons 53. (a) $8.99 \times 10^9\ \text{N}$; (b) $8.99\ \text{kN}$ 55. (a) 0.5; (b) 0.15;
(c) 0.85 57. $1.7 \times 10^8\ \text{N}$ 59. $-1.32 \times 10^{13}\ \text{C}$ 61. (a) $(0.829\ \text{N})\hat{i}$;
(b) $(-0.621\ \text{N})\hat{j}$ 63. $2.2 \times 10^{-6}\ \text{kg}$ 65. $4.68 \times 10^{-19}\ \text{N}$
67. (a) $2.72L$; (b) 0 69. (a) $5.1 \times 10^2\ \text{N}$; (b) $7.7 \times 10^{28}\ \text{m/s}^2$

CHAPTER 22

- CP** 1. (a) rightward; (b) leftward; (c) leftward; (d) rightward
(p and e have same charge magnitude, and p is farther)
2. (a) toward positive y ; (b) toward positive x ; (c) toward negative y
3. (a) leftward; (b) leftward; (c) decrease 4. (a) all tie; (b) 1 and 3
tie, then 2 and 4 tie
- Q** 1. a, b, c 3. (a) yes; (b) toward; (c) no (the field vectors are not
along the same line); (d) cancel; (e) add; (f) adding components;
(g) toward negative y 5. (a) to their left; (b) no 7. (a) 4, 3, 1, 2;
(b) 3, then 1 and 4 tie, then 2 9. a, b, c 11. e, b , then a and c tie,
then d (zero)
- P** 3. (a) $3.07 \times 10^{21}\ \text{N/C}$; (b) outward 5. $56\ \text{pC}$ 7. $(1.02 \times 10^5\ \text{N/C})\hat{j}$ 9. (a) $1.38 \times 10^{-10}\ \text{N/C}$; (b) 180° 11. $-30\ \text{cm}$
13. (a) $3.60 \times 10^{-6}\ \text{N/C}$; (b) $2.55 \times 10^{-6}\ \text{N/C}$; (c) $3.60 \times 10^{-4}\ \text{N/C}$;
(d) $7.09 \times 10^{-7}\ \text{N/C}$; (e) As the proton nears the disk, the forces on it from electrons e_s more nearly cancel. 15. (a) $160\ \text{N/C}$; (b) 45°
17. (a) -90° ; (b) $+2.0\ \mu\text{C}$; (c) $-1.6\ \mu\text{C}$ 19. (a) $qd/4\pi\epsilon_0 r^3$; (b) -90°
23. 0.506 25. (a) $1.62 \times 10^6\ \text{N/C}$; (b) -45° 27. (a) $23.8\ \text{N/C}$;
(b) -90° 29. 1.57 31. (a) $-5.19 \times 10^{-14}\ \text{C/m}$; (b) $1.57 \times 10^{-3}\ \text{N/C}$;
(c) -180° ; (d) $1.52 \times 10^{-8}\ \text{N/C}$; (e) $1.52 \times 10^{-8}\ \text{N/C}$ 35. $0.346\ \text{m}$
37. 28% 39. $-5e$ 41. (a) $1.5 \times 10^3\ \text{N/C}$; (b) $2.4 \times 10^{-16}\ \text{N}$; (c) up;
(d) $1.6 \times 10^{-26}\ \text{N}$; (e) 1.5×10^{10} 43. $3.51 \times 10^{15}\ \text{m/s}^2$
45. $6.6 \times 10^{-15}\ \text{N}$ 47. (a) $1.92 \times 10^{12}\ \text{m/s}^2$; (b) $1.96 \times 10^5\ \text{m/s}$
49. (a) $0.245\ \text{N}$; (b) -11.3° ; (c) $108\ \text{m}$; (d) $-21.6\ \text{m}$ 51. $2.6 \times 10^{-10}\ \text{N}$;
(b) $3.1 \times 10^{-8}\ \text{N}$; (c) moves to stigma 53. $27\ \mu\text{m}$ 55. (a) 2.7×10^6
 m/s ; (b) $1.0\ \text{kN/C}$ 57. (a) $9.30 \times 10^{-15}\ \text{C}\cdot\text{m}$; (b) $2.05 \times 10^{-11}\ \text{J}$
59. $1.22 \times 10^{-23}\ \text{J}$ 61. $(1/2\pi)(pE/I)^{0.5}$ 63. (a) $8.87 \times 10^{-15}\ \text{N}$;
(b) 120 65. 217° 67. $61\ \text{N/C}$ 69. (a) $47\ \text{N/C}$; (b) $27\ \text{N/C}$
71. $38\ \text{N/C}$ 73. (a) $-1.0\ \text{cm}$; (b) 0; (c) $10\ \text{pC}$ 75. $+1.00\ \mu\text{C}$
77. (a) $6.0\ \text{mm}$; (b) 180° 79. 9:30 81. (a) $-0.029\ \text{C}$; (b) repulsive
forces would explode the sphere 83. (a) $-1.49 \times 10^{-26}\ \text{J}$;
(b) $(-1.98 \times 10^{-26}\ \text{N}\cdot\text{m})\hat{k}$; (c) $3.47 \times 10^{-26}\ \text{J}$ 85. (a) top row: 4, 8,
12; middle row: 5, 10, 14; bottom row: 7, 11, 16; (b) $1.63 \times 10^{-19}\ \text{C}$
87. (a) $(-1.80\ \text{N/C})\hat{i}$; (b) $(43.2\ \text{N/C})\hat{i}$; (c) $(-6.29\ \text{N/C})\hat{i}$

CHAPTER 23

- CP** 1. (a) $+EA$; (b) $-EA$; (c) 0; (d) 0 2. (a) 2; (b) 3; (c) 1
3. (a) equal; (b) equal; (c) equal 4. 3 and 4 tie, then 2, 1
- Q** 1. (a) $8\ \text{N}\cdot\text{m}^2/\text{C}$; (b) 0 3. all tie 5. all tie 7. a, c , then b and d
tie (zero) 9. (a) 2, 1, 3; (b) all tie ($+4q$)
- P** 1. $-0.015\ \text{N}\cdot\text{m}^2/\text{C}$ 3. (a) 0; (b) $-3.92\ \text{N}\cdot\text{m}^2/\text{C}$; (c) 0; (d) 0
5. $3.01\ \text{nN}\cdot\text{m}^2/\text{C}$ 7. $2.0 \times 10^5\ \text{N}\cdot\text{m}^2/\text{C}$ 9. (a) $8.23\ \text{N}\cdot\text{m}^2/\text{C}$;
(b) $72.9\ \text{pC}$; (c) $8.23\ \text{N}\cdot\text{m}^2/\text{C}$; (d) $72.9\ \text{pC}$ 11. $-1.70\ \text{nC}$
13. $3.54\ \mu\text{C}$ 15. (a) 0; (b) 0.0417 17. (a) $37\ \mu\text{C}$; (b) $4.1 \times 10^6\ \text{N}\cdot\text{m}^2/\text{C}$
19. (a) $4.5 \times 10^{-7}\ \text{C}/\text{m}^2$; (b) $5.1 \times 10^4\ \text{N/C}$ 21. (a) $-3.0 \times 10^{-6}\ \text{C}$;
(b) $+1.3 \times 10^{-5}\ \text{C}$ 23. (a) $0.32\ \mu\text{C}$; (b) $0.14\ \mu\text{C}$ 25. $5.0\ \mu\text{C}/\text{m}$
27. $3.8 \times 10^{-8}\ \text{C}/\text{m}^2$ 29. (a) $0.214\ \text{N/C}$; (b) inward; (c) $0.855\ \text{N/C}$;
(d) outward; (e) $-3.40 \times 10^{-12}\ \text{C}$; (f) $-3.40 \times 10^{-12}\ \text{C}$ 31. (a) $2.3 \times 10^6\ \text{N/C}$; (b) outward; (c) $4.5 \times 10^5\ \text{N/C}$; (d) inward 33. (a) 0;
(b) 0; (c) $(-7.91 \times 10^{-11}\ \text{N/C})\hat{i}$ 35. -1.5 37. (a) $5.3 \times 10^7\ \text{N/C}$;
(b) $60\ \text{N/C}$ 39. $5.0\ \text{nC}/\text{m}^2$ 41. $0.44\ \text{mm}$ 43. (a) 0; (b) $1.31\ \mu\text{N/C}$;
(c) $3.08\ \mu\text{N/C}$; (d) $3.08\ \mu\text{N/C}$ 45. (a) $2.50 \times 10^4\ \text{N/C}$; (b) $1.35 \times 10^4\ \text{N/C}$ 47. $-7.5\ \text{nC}$ 49. (a) 0; (b) $56.2\ \text{mN/C}$; (c) $112\ \text{mN/C}$;
(d) $49.9\ \text{mN/C}$; (e) 0; (f) 0; (g) $-5.00\ \text{fC}$; (h) 0 51. $1.79 \times 10^{-11}\ \text{C}/\text{m}^2$
53. (a) $7.78\ \text{fC}$; (b) 0; (c) $5.58\ \text{mN/C}$; (d) $22.3\ \text{mN/C}$ 55. $6K\epsilon_0 r^3$
57. (a) 0; (b) $2.88 \times 10^4\ \text{N/C}$; (c) $200\ \text{N/C}$ 59. (a) $5.4\ \text{N/C}$;
(b) $6.8\ \text{N/C}$ 61. (a) 0; (b) $q_a/4\pi\epsilon_0 r^2$; (c) $(q_a + q_b)/4\pi\epsilon_0 r^2$
63. $-1.04\ \text{nC}$ 65. (a) 0.125; (b) 0.500 67. (a) $+2.0\ \text{nC}$;
(b) $-1.2\ \text{nC}$; (c) $+1.2\ \text{nC}$; (d) $+0.80\ \text{nC}$ 69. $(5.65 \times 10^4\ \text{N/C})\hat{j}$
71. (a) $-2.53 \times 10^{-2}\ \text{N}\cdot\text{m}^2/\text{C}$; (b) $+2.53 \times 10^{-2}\ \text{N}\cdot\text{m}^2/\text{C}$
75. $3.6\ \text{nC}$ 77. (a) $+4.0\ \mu\text{C}$; (b) $-4.0\ \mu\text{C}$ 79. (a) $693\ \text{kg/s}$;
(b) $693\ \text{kg/s}$; (c) $347\ \text{kg/s}$; (d) $347\ \text{kg/s}$; (e) $575\ \text{kg/s}$ 81. (a) $0.25R$;
(b) $2.0R$
- CHAPTER 24**
- CP** 1. (a) negative; (b) increase 2. (a) positive; (b) higher
3. (a) rightward; (b) 1, 2, 3, 5: positive; 4, negative; (c) 3, then 1, 2,
and 5 tie, then 4 4. all tie 5. a, c (zero), b 6. (a) 2, then 1 and 3
tie; (b) 3; (c) accelerate leftward
- Q** 1. $-4q/4\pi\epsilon_0 d$ 3. (a) 1 and 2; (b) none; (c) no; (d) 1 and 2, yes;
3 and 4, no 5. (a) higher; (b) positive; (c) negative; (d) all tie
7. (a) 0; (b) 0; (c) 0; (d) all three quantities still 0 9. (a) 3 and 4 tie,
then 1 and 2 tie; (b) 1 and 2, increase; 3 and 4, decrease
- P** 1. (a) $3.0 \times 10^5\ \text{C}$; (b) $3.6 \times 10^6\ \text{J}$ 3. 2.8×10^5 5. $8.8\ \text{mm}$
7. $-32.0\ \text{V}$ 9. (a) $1.87 \times 10^{-21}\ \text{J}$; (b) $-11.7\ \text{mV}$ 11. (a) $-0.268\ \text{mV}$;
(b) $-0.681\ \text{mV}$ 13. (a) $3.3\ \text{nC}$; (b) $12\ \text{nC}/\text{m}^2$ 15. (a) $0.54\ \text{mm}$;
(b) $790\ \text{V}$ 17. $0.562\ \text{mV}$ 19. (a) $6.0\ \text{cm}$; (b) $-12.0\ \text{cm}$ 21. $16.3\ \mu\text{V}$
23. (a) $24.3\ \text{mV}$; (b) 0 25. (a) $-2.30\ \text{V}$; (b) $-1.78\ \text{V}$ 27. $13\ \text{kV}$
29. $32.4\ \text{mV}$ 31. $47.1\ \mu\text{V}$ 33. $18.6\ \text{mV}$ 35. $(-12\ \text{V}/\text{m})\hat{i} + (12\ \text{V}/\text{m})\hat{j}$
37. $150\ \text{N/C}$ 39. $(-4.0 \times 10^{-16}\ \text{N})\hat{i} + (1.6 \times 10^{-16}\ \text{N})\hat{j}$
41. (a) $0.90\ \text{J}$; (b) $4.5\ \text{J}$ 43. $-0.192\ \text{pJ}$ 45. $2.5\ \text{km/s}$ 47. $22\ \text{km/s}$
49. $0.32\ \text{km/s}$ 51. (a) $+6.0 \times 10^4\ \text{V}$; (b) $-7.8 \times 10^5\ \text{V}$; (c) $2.5\ \text{J}$;
(d) increase; (e) same; (f) same 53. (a) $0.225\ \text{J}$; (b) $A 45.0\ \text{m/s}^2$,
 $B 22.5\ \text{m/s}^2$; (c) $A 7.75\ \text{m/s}$, $B 3.87\ \text{m/s}$ 55. $1.6 \times 10^{-9}\ \text{m}$
57. (a) $3.0\ \text{J}$; (b) $-8.5\ \text{m}$ 59. (a) proton; (b) $65.3\ \text{km/s}$ 61. (a) 12;
(b) 2 63. (a) $-1.8 \times 10^2\ \text{V}$; (b) $2.9\ \text{kV}$; (c) $-8.9\ \text{kV}$
65. $2.5 \times 10^{-8}\ \text{C}$ 67. (a) $12\ \text{kN/C}$; (b) $1.8\ \text{kV}$; (c) $5.8\ \text{cm}$
69. (a) $64\ \text{N/C}$; (b) $2.9\ \text{V}$; (c) 0 71. $p/2\pi\epsilon_0 r^3$ 73. (a) $3.6 \times 10^5\ \text{V}$;
(b) no 75. $6.4 \times 10^8\ \text{V}$ 77. $2.90\ \text{kV}$ 79. $7.0 \times 10^5\ \text{m/s}$
81. (a) $1.8\ \text{cm}$; (b) $8.4 \times 10^5\ \text{m/s}$; (c) $2.1 \times 10^{-17}\ \text{N}$; (d) positive;
(e) $1.6 \times 10^{-17}\ \text{N}$; (f) negative 83. (a) $+7.19 \times 10^{-10}\ \text{V}$;
(b) $+2.30 \times 10^{-28}\ \text{J}$; (c) $+2.43 \times 10^{-29}\ \text{J}$ 85. $2.30 \times 10^{-28}\ \text{J}$
87. 2.1 days 89. $2.30 \times 10^{-22}\ \text{J}$ 91. $1.48 \times 10^7\ \text{m/s}$ 93. $-1.92\ \text{MV}$

AN-8 ANSWERS

- 95.** (a) $Q/4\pi\epsilon_0 r$; (b) $(\rho/3\epsilon_0)(1.5r_2^2 - 0.50r^2 - r_1^3r^{-1})$,
 $\rho = Q/[(4\pi/3)(r_2^3 - r_1^3)]$; (c) $(\rho/2\epsilon_0)(r_2^2 - r_1^2)$, with ρ as in (b); (d) yes
101. (a) 0.484 MeV; (b) 0 **103.** -1.7 **105.** (a) 38 s; (b) 280 days

CHAPTER 25

- CP** **1.** (a) same; (b) same **2.** (a) decreases; (b) increases;
(c) decreases **3.** (a) $V, q/2$; (b) $V/2, q$
- Q** **1.** $a, 2; b, 1; c, 3$ **3.** (a) no; (b) yes; (c) all tie **5.** (a) same;
(b) same; (c) more; (d) more **7.** a , series; b , parallel; c , parallel
9. (a) increase; (b) same; (c) increase; (d) increase; (e) increase;
(f) increase **11.** parallel, C_1 alone, C_2 alone, series
- P** **1.** (a) 3.5 pF; (b) 3.5 pF; (c) 57 V **3.** (a) 144 pF; (b) 17.3 nC
5. 0.280 pF **7.** 6.79×10^{-4} F/m² **9.** 315 mC **11.** 3.16 μ F
13. 43 pF **15.** (a) 3.00 μ F; (b) 60.0 μ C; (c) 10.0 V; (d) 30.0 μ C;
(e) 10.0 V; (f) 20.0 μ C; (g) 5.00 V; (h) 20.0 μ C **17.** (a) 789 μ C;
(b) 78.9 V **19.** (a) 4.0 μ F; (b) 2.0 μ F **21.** (a) 50 V; (b) 5.0×10^{-5} C;
(c) 1.5×10^{-4} C **23.** (a) 4.5×10^{14} ; (b) 1.5×10^{14} ; (c) 3.0×10^{14} ; (d)
 4.5×10^{14} ; (e) up; (f) up **25.** 3.6 pC **27.** (a) 9.00 μ C;
(b) 16.0 μ C; (c) 9.00 μ C; (d) 16.0 μ C; (e) 8.40 μ C; (f) 16.8 μ C;
(g) 10.8 μ C; (h) 14.4 μ C **29.** 72 F **31.** 0.27 J **33.** 0.11 J/m³
35. (a) 9.16×10^{-18} J/m³; (b) 9.16×10^{-6} J/m³; (c) 9.16×10^6 J/m³;
(d) 9.16×10^{18} J/m³; (e) ∞ **37.** (a) 16.0 V; (b) 45.1 pJ; (c) 120 pJ; (d)
75.2 pJ **39.** (a) 190 V; (b) 95 mJ **41.** 81 pF/m **43.** Pyrex
45. 66 μ J **47.** 0.63 m^2 **49.** 17.3 pF **51.** (a) 10 kV/m; (b) 5.0 nC;
(c) 4.1 nC **53.** (a) 89 pF; (b) 0.12 nF; (c) 11 nC; (d) 11 nC;
(e) 10 kV/m; (f) 2.1 kV/m; (g) 88 V; (h) -0.17μ J **55.** (a) 0.107 nF;
(b) 7.79 nC; (c) 7.45 nC **57.** 45 μ C **59.** 16 μ C **61.** (a) 7.20 μ C;
(b) 18.0 μ C; (c) Battery supplies charges only to plates to which it is
connected; charges on other plates are due to electron transfers
between plates, in accord with new distribution of voltages across
the capacitors. So the battery does not directly supply charge on
capacitor 4. **63.** (a) 10 μ C; (b) 20 μ C **65.** 1.06 nC **67.** (a) 2.40 μ F;
(b) 0.480 mC; (c) 80 V; (d) 0.480 mC; (e) 120 V **69.** 4.9%
71. (a) 0.708 pF; (b) 0.600 ; (c) 1.02×10^{-9} J; (d) sucked in **73.** 5.3 V
75. 40 μ F **77.** (a) 200 kV/m; (b) 200 kV/m; (c) 1.77 μ C/m²;
(d) 4.60 μ C/m²; (e) -2.83μ C/m²

CHAPTER 26

- CP** **1.** 8 A, rightward **2.** (a)–(c) rightward **3.** a and c tie, then b
4. device 2 **5.** (a) and (b) tie, then (d), then (c)
- Q** **1.** tie of A , B , and C , then tie of $A+B$ and $B+C$, then
 $A+B+C$ **3.** (a) top-bottom, front-back, left-right; (b) top-
bottom, front-back, left-right; (c) top-bottom, front-back, left-right;
(d) top-bottom, front-back, left-right **5.** a, b , and c all tie, then d
7. (a) B, A, C ; (b) B, A, C **9.** (a) C, B, A ; (b) all tie; (c) A, B, C ;
(d) all tie
- P** **1.** (a) 1.2 kC; (b) 7.5×10^{21} **3.** 6.7 μ C/m² **5.** (a) 6.4 A/m²;
(b) north; (c) cross-sectional area **7.** 0.38 mm **9.** 18.1 μ A
11. (a) 1.33 A; (b) 0.666 A; (c) J_a **13.** 13 min **15.** 2.4 Ω
17. 2.0×10^6 ($\Omega \cdot \text{m}$)⁻¹ **19.** 2.0×10^{-8} $\Omega \cdot \text{m}$ **21.** (1.8×10^3) °C
23. 8.2×10^{-4} $\Omega \cdot \text{m}$ **25.** 54 Ω **27.** 3.0 **29.** 3.35×10^{-7} C
31. (a) 6.00 mA; (b) 1.59×10^{-8} V; (c) 21.2 n Ω **33.** (a) 38.3 mA;
(b) 109 A/m²; (c) 1.28 cm/s; (d) 227 V/m **35.** 981 k Ω **39.** 150 s
41. (a) 1.0 kW; (b) US\$0.25 **43.** 0.135 W **45.** (a) 10.9 A;
(b) 10.6 Ω ; (c) 4.50 MJ **47.** (a) 5.85 m; (b) 10.4 m **49.** (a) US\$4.46;
(b) 144 Ω ; (c) 0.833 A **51.** (a) 5.1 V; (b) 10 V; (c) 10 W; (d) 20 W
53. (a) 28.8 Ω ; (b) 2.60×10^{19} s⁻¹ **55.** 660 W **57.** 28.8 kC
59. (a) silver; (b) 51.6 n Ω **61.** (a) 2.3×10^{12} ; (b) 5.0×10^3 ; (c) 10 MV
63. 2.4 kW **65.** (a) 1.37; (b) 0.730 **67.** (a) -8.6%; (b) smaller
69. 146 kJ **71.** (a) 250°C; (b) yes **73.** 3.0×10^6 J/kg **75.** 560 W

CHAPTER 27

- CP** **1.** (a) rightward; (b) all tie; (c) b , then a and c tie;
(d) b , then a and c tie **2.** (a) all tie; (b) R_1, R_2, R_3 **3.** (a) less;
(b) greater; (c) equal **4.** (a) $V/2, i$; (b) $V, i/2$ **5.** (a) 1, 2, 4, 3;
(b) 4, tie of 1 and 2, then 3
- Q** **1.** (a) equal; (b) more **3.** parallel, R_2, R_1 , series **5.** (a) series;
(b) parallel; (c) parallel **7.** (a) less; (b) less; (c) more **9.** (a) same;
(b) same; (c) less; (d) more **11.** (a) all tie; (b) 1, 3, 2
- P** **1.** (a) 0.50 A; (b) 1.0 W; (c) 2.0 W; (d) 6.0 W; (e) 3.0 W; (f) sup-
plied; (g) absorbed **3.** (a) 14 V; (b) 1.0×10^2 W; (c) 6.0×10^2 W;
(d) 10 V; (e) 1.0×10^2 W **5.** 11 kJ **7.** (a) 80 J; (b) 67 J; (c) 13 J
9. 12.0 eV; (b) 6.53 W **11.** (a) 50 V; (b) 48 V; (c) negative
13. (a) 6.9 km; (b) 20Ω **15.** 8.0Ω **17.** (a) 0.004 Ω ; (b) 1
19. (a) 4.00Ω ; (b) parallel **21.** 5.56 A **23.** (a) 50 mA; (b) 60 mA;
(c) 9.0 V **25.** 3d **27.** 3.6×10^3 A **29.** (a) 0.333 A; (b) right;
(c) 720 J **31.** (a) -11 V; (b) -9.0 V **33.** 48.3 V **35.** (a) 5.25 V;
(b) 1.50 V; (c) 5.25 V; (d) 6.75 V **37.** 1.43 Ω **39.** (a) 0.150 Ω ;
(b) 240 W **41.** (a) 0.709 W; (b) 0.050 W; (c) 0.346 W; (d) 1.26 W;
(e) -0.158 W **43.** 9 **45.** (a) 0.67 A; (b) down; (c) 0.33 A; (d) up;
(e) 0.33 A; (f) up; (g) 3.3 V **47.** (a) 1.11 A; (b) 0.893 A; (c) 126 m
49. (a) 0.45 A **51.** (a) 55.2 mA; (b) 4.86 V; (c) 88.0 Ω ; (d) decrease
53. -3.0% **57.** 0.208 ms **59.** 4.61 **61.** (a) 2.41 μ s; (b) 161 pF
63. (a) 1.1 mA; (b) 0.55 mA; (c) 0.55 mA; (d) 0.82 mA; (e) 0.82 mA;
(f) 0; (g) 4.0×10^2 V; (h) 6.0×10^2 V **65.** 411 μ A **67.** 0.72 M Ω
69. (a) 0.955 μ C/s; (b) 1.08 μ W; (c) 2.74 μ W; (d) 3.82 μ W
71. (a) 3.00 A; (b) 3.75 A; (c) 3.94 A **73.** (a) 1.32×10^7 A/m²;
(b) 8.90 V; (c) copper; (d) 1.32×10^7 A/m²; (e) 51.1 V; (f) iron
75. (a) 3.0 kV; (b) 10 s; (c) 11 G Ω **77.** (a) 85.0 Ω ; (b) 915 Ω
81. 4.0 V **83.** (a) 24.8 Ω ; (b) 14.9 k Ω **85.** the cable **87.** -13 μ C
89. 20 Ω **91.** (a) 3.00 A; (b) down; (c) 1.60 A; (d) down; (e) supply;
(f) 55.2 W; (g) supply; (h) 6.40 W **93.** (a) 1.0 V; (b) 50 m Ω
95. 3 **99.** (a) 1.5 mA; (b) 0; (c) 1.0 mA

CHAPTER 28

- CP** **1.** $a, +z; b, -x; c, \vec{F}_B = 0$ **2.** (a) 2, then tie of 1 and 3 (zero);
(b) 4 **3.** (a) electron; (b) clockwise **4.** -y **5.** (a) all tie; (b) 1 and
4 tie, then 2 and 3 tie
- Q** **1.** (a) no, because \vec{v} and \vec{F}_B must be perpendicular; (b) yes;
(c) no, because \vec{B} and \vec{F}_B must be perpendicular
3. (a) +z and -z tie, then +y and -y tie, then +x and -x tie (zero);
(b) +y **5.** (a) \vec{F}_E ; (b) \vec{F}_B **7.** (a) \vec{B}_1 ; (b) \vec{B}_1 into page, \vec{B}_2 out of page;
(c) less **9.** (a) positive; (b) 2 → 1 and 2 → 4 tie, then 2 → 3 (which is
zero) **11.** (a) negative; (b) equal; (c) equal; (d) half-circle
- P** **1.** (a) 400 km/s; (b) 835 eV **3.** (a) $(6.2 \times 10^{-14}$ N) \hat{k} ;
(b) $(-6.2 \times 10^{-14}$ N) \hat{k} **5.** -2.0 T **7.** $(-11.4 \text{ V/m})\hat{i} - (6.00 \text{ V/m})\hat{j} +$
 $(4.80 \text{ V/m})\hat{k}$ **9.** $-(0.267 \text{ mT})\hat{k}$ **11.** 0.68 MV/m **13.** 7.4 μ V
15. (a) $(-600 \text{ mV/m})\hat{k}$; (b) 1.20 V **17.** (a) 2.60×10^6 m/s;
(b) 0.109 μ s; (c) 0.140 MeV; (d) 70.0 kV **19.** 1.2×10^{-9} kg/C
21. (a) 2.05×10^7 m/s; (b) 467 μ T; (c) 13.1 MHz; (d) 76.3 ns
23. 21.1 μ T **25.** (a) 0.978 MHz; (b) 96.4 cm **27.** (a) 495 mT;
(b) 22.7 mA; (c) 8.17 MJ **29.** 65.3 km/s **31.** 5.07 ns
33. (a) 0.358 ns; (b) 0.166 mm; (c) 1.51 mm **35.** (a) 200 eV;
(b) 20.0 keV; (c) 0.499% **37.** 2.4×10^2 m **39.** (a) 28.2 N;
(b) horizontally west **41.** (a) 467 mA; (b) right **43.** (a) 0; (b) 0.138 N;
(c) 0.138 N; (d) 0 **45.** $(-2.50 \text{ mN})\hat{j} + (0.750 \text{ mN})\hat{k}$ **47.** (a) 0.12 T;
(b) 31° **49.** $(-4.3 \times 10^{-3}$ N·m) \hat{i} **51.** 2.45 A **55.** (a) 2.86 A·m²;
(b) 1.10 A·m² **57.** (a) 12.7 A; (b) 0.0805 N·m **59.** (a) 0.30 A·m²;
(b) 0.024 N·m **61.** (a) -72.0 μ J; (b) $(96.0^\circ + 48.0^\circ)\mu\text{N}\cdot\text{m}$
63. (a) $(-9.7 \times 10^{-4}$ N·m) \hat{i} **65.** (a) 90°; (b) 1; (c) 1.28×10^{-7} N·m

67. (a) 20 min; (b) 5.9×10^{-2} N·m 69. 8.2 mm 71. 127 u
 73. (a) 6.3×10^{14} m/s²; (b) 3.0 mm 75. (a) 1.4; (b) 1.0
 77. (-500 V/m) \hat{j} 79. (a) 0.50; (b) 0.50; (c) 14 cm; (d) 14 cm
 81. ($0.80\hat{j} - 1.1\hat{k}$) mN 83. -40 mC 85. (a) $(12.8\hat{i} + 6.41\hat{j}) \times 10^{-22}$ N; (b) 90° ; (c) 173°

CHAPTER 29

- CP** 1. b, c, a 2. d, tie of a and c, then b 3. d, a, tie of b and c (zero)
Q 1. c, a, b 3. c, d, then a and b tie (zero) 5. a, c, b
 7. c and d tie, then b, a 9. b, a, d, c (zero) 11. (a) 1, 3, 2; (b) less
P 1. (a) $3.3 \mu\text{T}$; (b) yes 3. (a) 16 A ; (b) east 5. (a) 1.0 mT ;
 (b) out; (c) 0.80 mT ; (d) out 7. (a) $0.102 \mu\text{T}$; (b) out
 9. (a) opposite; (b) 30 A 11. (a) 4.3 A ; (b) out 13. 50.3 nT
 15. (a) $1.7 \mu\text{T}$; (b) into; (c) $6.7 \mu\text{T}$; (d) into 17. 132 nT 19. $5.0 \mu\text{T}$
 21. 256 nT 23. $(-7.75 \times 10^{-23} \text{ N})\hat{j}$ 25. 2.00 rad 27. 61.3 mA
 29. $(80 \mu\text{T})\hat{j}$ 31. (a) $20 \mu\text{T}$; (b) into 33. $(22.3 \text{ pT})\hat{j}$ 35. 88.4 pN/m
 37. $(-125 \mu\text{N/m})\hat{i} + (41.7 \mu\text{N/m})\hat{j}$ 39. 800 nN/m
 41. $(3.20 \text{ mN})\hat{j}$ 43. (a) 0; (b) 0.850 mT ; (c) 1.70 mT ; (d) 0.850 mT
 45. (a) $-2.5 \mu\text{T}\cdot\text{m}$; (b) 0 47. (a) 0; (b) $0.10 \mu\text{T}$; (c) $0.40 \mu\text{T}$
 49. (a) $533 \mu\text{T}$; (b) $400 \mu\text{T}$ 51. 0.30 mT 53. 0.272 A
 55. (a) 4.77 cm ; (b) $35.5 \mu\text{T}$ 57. (a) $2.4 \text{ A}\cdot\text{m}^2$; (b) 46 cm
 59. $0.47 \text{ A}\cdot\text{m}^2$ 61. (a) $79 \mu\text{T}$; (b) $1.1 \times 10^{-6} \text{ N}\cdot\text{m}$ 63. (a) $(0.060 \text{ A}\cdot\text{m}^2)\hat{j}$;
 (b) $(96 \text{ pT})\hat{j}$ 65. 1.28 mm 69. (a) 15 A ; (b) $-z$ 71. 7.7 mT
 73. (a) $15.3 \mu\text{T}$ 75. (a) $(0.24\hat{i}) \text{ nT}$; (b) 0; (c) $(-43\hat{k}) \text{ pT}$; (d) $(0.14\hat{k}) \text{ nT}$
 79. (a) 4.8 mT ; (b) 0.93 mT ; (c) 0 83. $(-0.20 \text{ mT})\hat{k}$
 87. (a) $\mu_0 ir/2\pi r^2$; (b) $\mu_0 i/2\pi r$; (c) $\mu_0 i(a^2 - r^2)/2\pi(a^2 - b^2)r$; (d) 0

CHAPTER 30

- CP** 1. b, then d and e tie, and then a and c tie (zero) 2. a and b tie, then c (zero) 3. c and d tie, then a and b tie 4. b, out; c, out; d, into; e, into 5. d and e 6. (a) 2, 3, 1 (zero); (b) 2, 3, 1
 7. a and b tie, then c
Q 1. out 3. (a) all tie (zero); (b) 2, then 1 and 3 tie (zero) 5. d and c tie, then b, a 7. (a) more; (b) same; (c) same; (d) same (zero)
 9. (a) all tie (zero); (b) 1 and 2 tie, then 3; (c) all tie (zero)
P 1. 0 3. 30 mA 5. 0 7. (a) 31 mV ; (b) left 9. 0.198 mV
 11. (b) 0.796 m^2 13. 29.5 mC 15. (a) 21.7 V ; (b) counterclockwise 17. (a) $1.26 \times 10^{-4} \text{ T}$; (b) 0; (c) $1.26 \times 10^{-4} \text{ T}$; (d) yes;
 (e) $5.04 \times 10^{-8} \text{ V}$ 19. 5.50 kV 21. (a) 40 Hz ; (b) 3.2 mV
 23. (a) $\mu_0 iR^2\pi r^2/2x^3$; (b) $3\mu_0 i\pi R^2r^2v/2x^4$; (c) counterclockwise
 25. (a) $13 \mu\text{Wb}/\text{m}$; (b) 17%; (c) 0 27. (a) $80 \mu\text{V}$; (b) clockwise
 29. (a) 48.1 mV ; (b) 2.67 mA ; (c) 0.129 mW 31. $3.68 \mu\text{W}$
 33. (a) $240 \mu\text{V}$; (b) 0.600 mA ; (c) $0.144 \mu\text{W}$; (d) $2.87 \times 10^{-8} \text{ N}$;
 (e) $0.144 \mu\text{W}$ 35. (a) 0.60 V ; (b) up; (c) 1.5 A ; (d) clockwise;
 (e) 0.90 W ; (f) 0.18 N ; (g) 0.90 W 37. (a) $71.5 \mu\text{V}/\text{m}$; (b) $143 \mu\text{V}/\text{m}$
 39. 0.15 V/m 41. (a) 2.45 mWb ; (b) 0.645 mH 43. $1.81 \mu\text{H}/\text{m}$
 45. (a) decreasing; (b) 0.68 mH 47. (b) $L_{eq} = \sum L_j$; sum from $j = 1$ to $j = N$ 49. 59.3 mH 51. 46Ω 53. (a) 8.45 ns ; (b) 7.37 mA
 55. 6.91 57. (a) 1.5 s 59. (a) $i[1 - \exp(-Rt/L)]$; (b) $(L/R) \ln 2$
 61. (a) 97.9 H ; (b) 0.196 mJ 63. 25.6 ms 65. (a) 18.7 J ; (b) 5.10 J ;
 (c) 13.6 J 67. (a) 34.2 J/m^3 ; (b) 49.4 mJ 69. $1.5 \times 10^8 \text{ V/m}$
 71. (a) 1.0 J/m^3 ; (b) $4.8 \times 10^{-15} \text{ J/m}^3$ 73. (a) 1.67 mH ; (b) 6.00 mWb
 75. $13 \mu\text{H}$ 77. (b) have the turns of the two solenoids wrapped in opposite directions 79. (a) 2.0 A ; (b) 0; (c) 2.0 A ; (d) 0; (e) 10 V ;
 (f) 2.0 A/s ; (g) 2.0 A ; (h) 1.0 A ; (i) 3.0 A ; (j) 10 V ; (k) 0; (l) 0
 81. (a) $10 \mu\text{T}$; (b) out; (c) $3.3 \mu\text{T}$; (d) out 83. 0.520 ms
 85. (a) $(4.4 \times 10^7 \text{ m/s}^2)\hat{i}$; (b) 0; (c) $(-4.4 \times 10^7 \text{ m/s}^2)\hat{i}$
 87. (a) 0.40 V ; (b) 20 A 89. (a) 10 A ; (b) $1.0 \times 10^2 \text{ J}$ 91. (a) 0;
 (b) $8.0 \times 10^2 \text{ A/s}$; (c) 1.8 mA ; (d) $4.4 \times 10^2 \text{ A/s}$; (e) 4.0 mA ; (f) 0
 93. 1.15 W 95. (a) 20 A/s ; (b) 0.75 A 97. 12 A/s

CHAPTER 31

- CP** 1. (a) $T/2$; (b) T ; (c) $T/2$; (d) $T/4$ 2. (a) 5 V ; (b) $150 \mu\text{J}$
 3. (a) remains the same; (b) remains the same 4. (a) C, B, A ; (b) 1,
 $A; 2, B; 3, S; 4, C$; (c) A 5. (a) remains the same; (b) increases;
 (c) remains the same; (d) decreases 6. (a) 1, lags; 2, leads; 3, in
 phase; (b) 3 ($\omega_d = \omega$ when $X_L = X_C$) 7. (a) increase (circuit is
 mainly capacitive; increase C to decrease X_C to be closer to reso-
 nance for maximum P_{avg}); (b) closer 8. (a) greater; (b) step-up
Q 1. b, a, c 3. (a) $T/4$; (b) $T/4$; (c) $T/2$; (d) $T/2$ 5. c, b, a 7. a
 inductor; b resistor; c capacitor 9. (a) positive; (b) decreased (to
 decrease X_L and get closer to resonance); (c) decreased (to increase
 X_C and get closer to resonance) 11. (a) rightward, increase (X_L
 increases, closer to resonance); (b) rightward, increase (X_C decreases,
 closer to resonance); (c) rightward, increase (ω_d/ω increases, closer to
 resonance)
P 1. (a) $1.17 \mu\text{J}$; (b) 5.58 mA 3. (a) $6.00 \mu\text{s}$; (b) 167 kHz ; (c) $3.00 \mu\text{s}$
 5. 45.2 mA 7. (a) 1.25 kg ; (b) 372 N/m ; (c) $1.75 \times 10^{-4} \text{ m}$;
 (d) 3.02 mm/s 9. $7.0 \times 10^{-4} \text{ s}$ 11. (a) 6.0 ; (b) 36 pF ; (c) 0.22 mH
 13. (a) 0.180 mC ; (b) $70.7 \mu\text{s}$; (c) 66.7 W 15. (a) 3.0 nC ; (b) 1.7 mA ;
 (c) 4.5 nJ 17. (a) 275 Hz ; (b) 365 mA 21. (a) $356 \mu\text{s}$; (b) 2.50 mH ;
 (c) 3.20 mJ 23. (a) $1.98 \mu\text{J}$; (b) $5.56 \mu\text{C}$; (c) 12.6 mA ; (d) -46.9° ;
 (e) $+46.9^\circ$ 25. $8.66 \text{ m}\Omega$ 29. (a) 95.5 mA ; (b) 11.9 mA
 31. (a) 0.65 kHz ; (b) 24Ω 33. (a) 6.73 ms ; (b) 11.2 ms ;
 (c) inductor; (d) 138 mH 35. 89Ω 37. 7.61 A 39. (a) 267Ω ;
 (b) -41.5° ; (c) 135 mA 41. (a) 206Ω ; (b) 13.7° ; (c) 175 mA
 43. (a) 218Ω ; (b) 23.4° ; (c) 165 mA 45. (a) yes; (b) 1.0 kV
 47. (a) 224 rad/s ; (b) 6.00 A ; (c) 219 rad/s ; (d) 228 rad/s ; (e) 0.040
 49. (a) 796 Hz ; (b) no change; (c) decreased; (d) increased
 53. (a) 12.1Ω ; (b) 1.19 kW 55. 1.84 A 57. (a) $117 \mu\text{F}$; (b) 0;
 (c) 90.0 W ; (d) 0° ; (e) 1; (f) 0; (g) -90° ; (h) 0 59. (a) 2.59 A ;
 (b) 38.8 V ; (c) 159 V ; (d) 224 V ; (e) 64.2 V ; (f) 75.0 V ; (g) 100 W ; (h) 0;
 (i) 0 61. (a) 0.743 ; (b) lead; (c) capacitive; (d) no; (e) yes; (f) no;
 (g) yes; (h) 33.4 W 63. (a) 2.4 V ; (b) 3.2 mA ; (c) 0.16 A
 65. (a) 1.9 V ; (b) 5.9 W ; (c) 19 V ; (d) $5.9 \times 10^2 \text{ W}$; (e) 0.19 kV ; (f) 59 kW
 67. (a) 6.73 ms ; (b) 2.24 ms ; (c) capacitor; (d) $59.0 \mu\text{F}$
 69. (a) -0.405 rad ; (b) 2.76 A ; (c) capacitive 71. (a) 64.0Ω ;
 (b) 50.9Ω ; (c) capacitive 73. (a) $2.41 \mu\text{H}$; (b) 21.4 pJ ; (c) 82.2 nC
 75. (a) 39.1Ω ; (b) 21.7Ω ; (c) capacitive 79. (a) $0.577 Q$; (b) 0.152
 81. (a) 45.0° ; (b) 70.7Ω 83. 1.84 kHz 85. (a) $0.689 \mu\text{H}$;
 (b) 17.9 pJ ; (c) $0.110 \mu\text{C}$ 87. (a) 165Ω ; (b) 313 mH ; (c) $14.9 \mu\text{F}$
- CHAPTER 32**
- CP** 1. d, b, c, a (zero) 2. a, c, b, d (zero) 3. tie of b, c, and d, then a
 4. (a) 2; (b) 1 5. (a) away; (b) away; (c) less 6. (a) toward;
 (b) toward; (c) less
Q 1. 1, 2, b, 3 c and d 3. a, decreasing; b, decreasing
 5. supplied 7. (a) a and b tie, then c, d; (b) none (because plate
 lacks circular symmetry, \vec{B} not tangent to any circular loop);
 (c) none 9. (a) 1 up, 2 up, 3 down; (b) 1 down, 2 up, 3 zero
 11. (a) 1, 3, 2; (b) 2
P 1. $+3 \text{ Wb}$ 3. (a) $47.4 \mu\text{Wb}$; (b) inward 5. $2.4 \times 10^{13} \text{ V/m}\cdot\text{s}$
 7. (a) $1.18 \times 10^{-19} \text{ T}$; (b) $1.06 \times 10^{-19} \text{ T}$ 9. (a) $5.01 \times 10^{-22} \text{ T}$;
 (b) $4.51 \times 10^{-22} \text{ T}$ 11. (a) 1.9 pT 13. $7.5 \times 10^5 \text{ V/s}$
 17. (a) 0.324 V/m ; (b) $2.87 \times 10^{-16} \text{ A}$; (c) 2.87×10^{-18}
 19. (a) 75.4 nT ; (b) 67.9 nT 21. (a) 27.9 nT ; (b) 15.1 nT
 23. (a) 2.0 A ; (b) $2.3 \times 10^{11} \text{ V/m}\cdot\text{s}$; (c) 0.50 A ; (d) $0.63 \mu\text{T}\cdot\text{m}$
 25. (a) $0.63 \mu\text{T}$; (b) $2.3 \times 10^{12} \text{ V/m}\cdot\text{s}$ 27. (a) 0.71 A ; (b) 0; (c) 2.8 A
 29. (a) $7.60 \mu\text{A}$; (b) $859 \text{ kV}\cdot\text{m/s}$; (c) 3.39 mm ; (d) 5.16 pT 31. $55 \mu\text{T}$
 33. (a) 0; (b) 0; (c) 0; (d) $\pm 3.2 \times 10^{-25} \text{ J}$; (e) $-3.2 \times 10^{-34} \text{ J}\cdot\text{s}$;
 (f) $2.8 \times 10^{-23} \text{ J/T}$; (g) $-9.7 \times 10^{-25} \text{ J}$; (h) $\pm 3.2 \times 10^{-25} \text{ J}$
 35. (a) $-9.3 \times 10^{-24} \text{ J/T}$; (b) $1.9 \times 10^{-23} \text{ J/T}$ 37. (b) $+x$;

AN-10 ANSWERS

- (c) clockwise; (d) $+x$ **39.** yes **41.** 20.8 mJ/T **43.** (b) K_i/B ; (c) $-z$; (d) 0.31 kA/m **47.** (a) $1.8 \times 10^2 \text{ km}$; (b) 2.3×10^{-5} **49.** (a) $3.0 \mu\text{T}$; (b) $5.6 \times 10^{-10} \text{ eV}$ **51.** $5.15 \times 10^{-24} \text{ A}\cdot\text{m}^2$ **53.** (a) 0.14 A ; (b) $79 \mu\text{C}$ **55.** (a) $6.3 \times 10^8 \text{ A}$; (b) yes; (c) no **57.** 0.84 kJ/T **59.** (a) $(1.2 \times 10^{-13} \text{ T}) \exp[-t/(0.012 \text{ s})]$; (b) $5.9 \times 10^{-15} \text{ T}$ **63.** (a) 27.5 mm ; (b) 110 mm **65.** 8.0 A **67.** (a) $-8.8 \times 10^{15} \text{ V/m}\cdot\text{s}$; (b) $5.9 \times 10^{-7} \text{ T}\cdot\text{m}$ **69.** (b) sign is minus; (c) no, because there is compensating positive flux through open end nearer to magnet **71.** (b) $-x$; (c) counterclockwise; (d) $-x$ **73.** (a) 7; (b) 7; (c) $3h/2\pi$; (d) $3eh/4\pi m$; (e) $3.5h/2\pi$; (f) 8 **75.** (a) 9; (b) $3.71 \times 10^{-23} \text{ J/T}$; (c) $+9.27 \times 10^{-24} \text{ J}$; (d) $-9.27 \times 10^{-24} \text{ J}$

CHAPTER 33

- CP** **1.** (a) (Use Fig. 33-5.) On right side of rectangle, \vec{E} is in negative y direction; on left side, $\vec{E} + d\vec{E}$ is greater and in same direction; (b) \vec{E} is downward. On right side, \vec{B} is in negative z direction; on left side, $\vec{B} + d\vec{B}$ is greater and in same direction. **2.** positive direction of x **3.** (a) same; (b) decrease **4.** a, d, b, c (zero) **5.** a **Q** **1.** (a) positive direction of z ; (b) x **3.** (a) same; (b) increase; (c) decrease **5.** (a) and (b) $A = 1, n = 4, \theta = 30^\circ$ **7.** a, b, c **9.** B **11.** none **P** **1.** 7.49 GHz **3.** (a) 515 nm ; (b) 610 nm ; (c) 555 nm ; (d) $5.41 \times 10^{14} \text{ Hz}$; (e) $1.85 \times 10^{-15} \text{ s}$ **5.** $5.0 \times 10^{-21} \text{ H}$ **7.** 1.2 MW/m^2 **9.** 0.10 MJ **11.** (a) 6.7 nT ; (b) y ; (c) negative direction of y **13.** (a) 1.03 kV/m ; (b) $3.43 \mu\text{T}$ **15.** (a) 87 mV/m ; (b) 0.29 nT ; (c) 6.3 kW **17.** (a) 6.7 nT ; (b) 5.3 mW/m^2 ; (c) 6.7 W **19.** $1.0 \times 10^7 \text{ Pa}$ **21.** $5.9 \times 10^{-8} \text{ Pa}$ **23.** (a) $4.68 \times 10^{11} \text{ W}$; (b) any chance disturbance could move sphere from directly above source—the two force vectors no longer along the same axis **27.** (a) $1.0 \times 10^8 \text{ Hz}$; (b) $6.3 \times 10^8 \text{ rad/s}$; (c) 2.1 m^{-1} ; (d) $1.0 \mu\text{T}$; (e) z ; (f) $1.2 \times 10^2 \text{ W/m}^2$; (g) $8.0 \times 10^{-7} \text{ N}$; (h) $4.0 \times 10^{-7} \text{ Pa}$ **29.** 1.9 mm/s **31.** (a) $0.17 \mu\text{m}$; (b) toward the Sun **33.** 3.1% **35.** 4.4 W/m^2 **37.** (a) 2 sheets; (b) 5 sheets **39.** (a) 1.9 V/m ; (b) $1.7 \times 10^{-11} \text{ Pa}$ **41.** 20° or 70° **43.** 0.67 **45.** 1.26 **47.** 1.48 **49.** 180° **51.** (a) 56.9° ; (b) 35.3° **55.** 1.07 m **57.** 182 cm **59.** (a) 48.9° ; (b) 29.0° **61.** (a) 26.8° ; (b) yes **63.** (a) $(1 + \sin^2 \theta)^{0.5}$; (b) $2^{0.5}$; (c) yes; (d) no **65.** 23.2° **67.** (a) 1.39 ; (b) 28.1° ; (c) no **69.** 49.0° **71.** (a) 0.50 ms ; (b) 8.4 min ; (c) 2.4 h ; (d) 5446 B.C. **73.** (a) $(16.7 \text{ nT}) \sin[(1.00 \times 10^6 \text{ m}^{-1})z + (3.00 \times 10^{14} \text{ s}^{-1})t]$; (b) $6.28 \mu\text{m}$; (c) 20.9 fs ; (d) 33.2 mW/m^2 ; (e) x ; (f) infrared **75.** 1.22 **77.** (c) 137.6° (d) 139.4° (e) 1.7° **81.** (a) z axis; (b) $7.5 \times 10^{14} \text{ Hz}$; (c) 1.9 kW/m^2 **83.** (a) white; (b) white dominated by red end; (c) no refracted light **85.** $1.5 \times 10^{-9} \text{ m/s}^2$ **87.** (a) $3.5 \mu\text{W/m}^2$; (b) $0.78 \mu\text{W}$; (c) $1.5 \times 10^{-17} \text{ W/m}^2$; (d) $1.1 \times 10^{-7} \text{ V/m}$; (e) 0.25 ft **89.** (a) 55.8° ; (b) 55.5° **91.** (a) 83 W/m^2 ; (b) 1.7 MW **93.** 35°

CHAPTER 34

- CP** **1.** $0.2d, 1.8d, 2.2d$ **2.** (a) real; (b) inverted; (c) same **3.** (a) e ; (b) virtual, same **4.** virtual, same as object, diverging **Q** **1.** (a) a ; (b) c **3.** (a) a and c ; (b) three times; (c) you **5.** convex **7.** (a) all but variation 2; (b) 1, 3, 4: right, inverted; 5, 6: left, same **9.** d (infinite), tie of a and b , then c **P** **1.** 9.10 m **3.** 1.11 **5.** 351 cm **7.** 10.5 cm **9.** (a) $+24 \text{ cm}$; (b) $+36 \text{ cm}$; (c) -2.0 ; (d) R ; (e) I ; (f) same **11.** (a) -20 cm ; (b) -4.4 cm ; (c) $+0.56$; (d) V ; (e) NI ; (f) opposite **13.** (a) $+36 \text{ cm}$; (b) -36 cm ; (c) $+3.0$; (d) V ; (e) NI ; (f) opposite **15.** (a) -16 cm ; (b) -4.4 cm ; (c) $+0.44$; (d) V ; (e) NI ; (f) opposite **17.** (b) plus; (c) $+40 \text{ cm}$; (e) -20 cm ; (f) $+2.0$; (g) V ; (h) NI ; (i) opposite **19.** (a) convex; (b) -20 cm ; (d) $+20 \text{ cm}$; (f) $+0.50$; (g) V ; (h) NI ;

- (i) opposite **21.** (a) concave; (c) $+40 \text{ cm}$; (e) $+60 \text{ cm}$; (f) -2.0 ; (g) R ; (h) I ; (i) same **23.** (a) convex; (b) minus; (c) -60 cm ; (d) $+1.2 \text{ m}$; (e) -24 cm ; (g) V ; (h) NI ; (i) opposite **25.** (a) concave; (b) $+8.6 \text{ cm}$; (c) $+17 \text{ cm}$; (e) $+12 \text{ cm}$; (f) minus; (g) R ; (i) same **27.** (a) convex; (c) -60 cm ; (d) $+30 \text{ cm}$; (f) $+0.50$; (g) V ; (h) NI ; (i) opposite **29.** (b) -20 cm ; (c) minus; (d) $+5.0 \text{ cm}$; (e) minus; (f) $+0.80$; (g) V ; (h) NI ; (i) opposite **31.** (b) 0.56 cm/s ; (c) 11 m/s ; (d) 6.7 cm/s **33.** (c) -33 cm ; (e) V ; (f) same **35.** (d) -26 cm ; (e) V ; (f) same **37.** (c) $+30 \text{ cm}$; (e) V ; (f) same **39.** (a) 2.00 ; (b) none **41.** (a) $+40 \text{ cm}$; (b) ∞ **43.** 5.0 mm **45.** 1.86 mm **47.** (a) 45 mm ; (b) 90 mm **49.** 22 cm **51.** (a) -48 cm ; (b) $+4.0$; (c) V ; (d) NI ; (e) same **53.** (a) -4.8 cm ; (b) $+0.60$; (c) V ; (d) NI ; (e) same **55.** (a) -8.6 cm ; (b) $+0.39$; (c) V ; (d) NI ; (e) same **57.** (a) $+36 \text{ cm}$; (b) -0.80 ; (c) R ; (d) I ; (e) opposite **59.** (a) $+55 \text{ cm}$; (b) -0.74 ; (c) R ; (d) I ; (e) opposite **61.** (a) -18 cm ; (b) $+0.76$; (c) V ; (d) NI ; (e) same **63.** (a) -30 cm ; (b) $+0.86$; (c) V ; (d) NI ; (e) same **65.** (a) -7.5 cm ; (b) $+0.75$; (c) V ; (d) NI ; (e) same **67.** (a) $+84 \text{ cm}$; (b) -1.4 ; (c) R ; (d) I ; (e) opposite **69.** (a) C ; (d) -10 cm ; (e) $+2.0$; (f) V ; (g) NI ; (h) same **71.** (a) D ; (b) -5.3 cm ; (d) -4.0 cm ; (f) V ; (g) NI ; (h) same **73.** (a) C ; (b) $+3.3 \text{ cm}$; (d) $+5.0 \text{ cm}$; (f) R ; (g) I ; (h) opposite **75.** (a) D ; (b) minus; (d) -3.3 cm ; (e) $+0.67$; (f) V ; (g) NI **77.** (a) C ; (b) $+80 \text{ cm}$; (d) -20 cm ; (f) V ; (g) NI ; (h) same **79.** (a) C ; (b) plus; (d) -13 cm ; (e) $+1.7$; (f) V ; (g) NI ; (h) same **81.** (a) $+24 \text{ cm}$; (b) $+6.0$; (c) R ; (d) NI ; (e) opposite **83.** (a) $+3.1 \text{ cm}$; (b) -0.31 ; (c) R ; (d) I ; (e) opposite **85.** (a) -4.6 cm ; (b) $+0.69$; (c) V ; (d) NI ; (e) same **87.** (a) -5.5 cm ; (b) $+0.12$; (c) V ; (d) NI ; (e) same **89.** (a) 13.0 cm ; (b) 5.23 cm ; (c) -3.25 ; (d) 3.13 ; (e) -10.2 **91.** (a) 2.35 cm ; (b) decrease **93.** (a) 3.5 ; (b) 2.5 **95.** (a) $+8.6 \text{ cm}$; (b) $+2.6$; (c) R ; (d) NI ; (e) opposite **97.** (a) $+7.5 \text{ cm}$; (b) -0.75 ; (c) R ; (d) I ; (e) opposite **99.** (a) $+24 \text{ cm}$; (b) -0.58 ; (c) R ; (d) I ; (e) opposite **105.** (a) 3.00 cm ; (b) 2.33 cm **107.** (a) 40 cm ; (b) 20 cm ; (c) -40 cm ; (d) 40 cm **109.** (a) 20 cm ; (b) 15 cm **111.** (a) 6.0 mm ; (b) 1.6 kW/m^2 ; (c) 4.0 cm

CHAPTER 35

- CP** **1.** b (least n), c, a **2.** (a) top; (b) bright intermediate illumination (phase difference is 2.1 wavelengths) **3.** (a) $3\lambda, 3$; (b) $2.5\lambda, 2.5$ **4.** a and d tie (amplitude of resultant wave is $4E_0$), then b and c tie (amplitude of resultant wave is $2E_0$) **5.** (a) 1 and 4; (b) 1 and 4 **Q** **1.** (a) decrease; (b) decrease; (c) decrease; (d) blue **3.** (a) $2d$; (b) (odd number) $\lambda/2$; (c) $\lambda/4$ **5.** (a) intermediate closer to maximum, $m = 2$; (b) minimum, $m = 3$; (c) intermediate closer to maximum, $m = 2$; (d) maximum, $m = 1$ **7.** (a) maximum; (b) minimum; (c) alternates **9.** (a) peak; (b) valley **11.** c, d **P** **1.** (a) 155 nm ; (b) 310 nm **3.** (a) $3.60 \mu\text{m}$; (b) intermediate closer to fully constructive **5.** $4.55 \times 10^7 \text{ m/s}$ **7.** 1.56 **9.** (a) $1.55 \mu\text{m}$; (b) $4.65 \mu\text{m}$ **11.** (a) 1.70 ; (b) 1.70 ; (c) 1.30 ; (d) all tie **13.** (a) 0.833 ; (b) intermediate closer to fully constructive **15.** 648 nm **17.** 16 **19.** 2.25 mm **21.** $72 \mu\text{m}$ **23.** 0 **25.** $7.88 \mu\text{m}$ **27.** $6.64 \mu\text{m}$ **29.** 2.65 **31.** $27 \sin(\omega t + 8.5^\circ)$ **33.** $(17.1 \mu\text{V/m}) \sin[(2.0 \times 10^{14} \text{ rad/s})t]$ **35.** 120 nm **37.** 70.0 nm **39.** (a) $0.117 \mu\text{m}$; (b) $0.352 \mu\text{m}$ **41.** 161 nm **43.** 560 nm **45.** 478 nm **47.** 509 nm **49.** 273 nm **51.** 409 nm **53.** 338 nm **55.** (a) 552 nm ; (b) 442 nm **57.** 608 nm **59.** 528 nm **61.** 455 nm **63.** 248 nm **65.** 339 nm **67.** 329 nm **69.** $1.89 \mu\text{m}$ **71.** 0.012° **73.** 140 **75.** $[(m + \frac{1}{2})\lambda R]^{0.5}$, for $m = 0, 1, 2, \dots$ **77.** 1.00 m **79.** 588 nm **81.** 1.00030 **83.** (a) 50.0 nm ; (b) 36.2 nm **85.** 0.23° **87.** (a) 1500 nm ; (b) 2250 nm ; (c) 0.80 **89.** $x = (D/2a)(m + 0.5)\lambda$, for $m = 0, 1, 2, \dots$ **91.** (a) 22° ; (b) refraction reduces θ **93.** 600 nm **95.** (a) $1.75 \mu\text{m}$; (b) 4.8 mm **97.** $I_m \cos^2(2\pi x/\lambda)$ **99.** (a) 42.0 ps ; (b) 42.3 ps ; (c) 43.2 ps ; (d) 41.8 ps ; (e) **101.** $33 \mu\text{m}$

CHAPTER 36

- CP** 1. (a) expand; (b) expand 2. (a) second side maximum; (b) 2.5 3. (a) red; (b) violet 4. diminish 5. (a) left; (b) less
- Q** 1. (a) $m = 5$ minimum; (b) (approximately) maximum between the $m = 4$ and $m = 5$ minima 3. (a) A, B, C ; (b) A, B, C 5. (a) 1 and 3 tie, then 2 and 4 tie; (b) 1 and 2 tie, then 3 and 4 tie 7. (a) larger; (b) red 9. (a) decrease; (b) same; (c) remain in place 11. (a) A ; (b) left; (c) left; (d) right
- P** 1. (a) 2.5 mm; (b) 2.2×10^{-4} rad 3. (a) 70 cm; (b) 1.0 mm 5. (a) 700 nm; (b) 4; (c) 6 7. 60.4 μm 9. 1.77 mm 11. 160° 13. (a) 0.18°; (b) 0.46 rad; (c) 0.93 15. (d) 52.5°; (e) 10.1°; (f) 5.06° 17. (b) 0; (c) -0.500; (d) 4.493 rad; (e) 0.930; (f) 7.725 rad; (g) 1.96 19. (a) 19 cm; (b) larger 21. (a) 1.1×10^4 km; (b) 11 km 23. (a) 1.3×10^{-4} rad; (b) 10 km 25. 50 m 27. 1.6×10^3 km 29. (a) 8.8×10^{-7} rad; (b) 8.4×10^7 km; (c) 0.025 mm 31. (a) 0.346°; (b) 0.97° 33. (a) 17.1 m; (b) 1.37×10^{-10} 35. 5 37. 3 39. (a) 5.0 μm ; (b) 20 μm 41. (a) 7.43×10^{-3} ; (b) between the $m = 6$ minimum (the seventh one) and the $m = 7$ maximum (the seventh side maximum); (c) between the $m = 3$ minimum (the third one) and the $m = 4$ minimum (the fourth one) 43. (a) 9; (b) 0.255 45. (a) 62.1°; (b) 45.0°; (c) 32.0° 47. 3 49. (a) 6.0 μm ; (b) 1.5 μm ; (c) 9; (d) 7; (e) 6 51. (a) 2.1°; (b) 21°; (c) 11 53. (a) 470 nm; (b) 560 nm 55. 3.65×10^3 57. (a) $0.032^\circ/\text{nm}$; (b) 4.0×10^4 ; (c) $0.076^\circ/\text{nm}$; (d) 8.0×10^4 ; (e) $0.24^\circ/\text{nm}$; (f) 1.2×10^5 59. 0.15 nm 61. (a) 10 μm ; (b) 3.3 mm 63. 1.09×10^3 rulings/mm 65. (a) 0.17 nm; (b) 0.13 nm 67. (a) 25 pm; (b) 38 pm 69. 0.26 nm 71. (a) 15.3°; (b) 30.6°; (c) 3.1°; (d) 37.8° 73. (a) $0.7071a_0$; (b) $0.4472a_0$; (c) $0.3162a_0$; (d) $0.2774a_0$; (e) $0.2425a_0$ 75. (a) 625 nm; (b) 500 nm; (c) 416 nm 77. 3.0 mm 83. (a) 13; (b) 6 85. 59.5 pm 87. 4.9 km 89. 1.36×10^4 91. 2 93. 4.7 cm 97. 36 cm 99. (a) fourth; (b) seventh

CHAPTER 37

- CP** 1. (a) same (speed of light postulate); (b) no (the start and end of the flight are spatially separated); (c) no (because his measurement is not a proper time) 2. (a) Eq. 2; (b) +0.90c; (c) 25 ns; (d) -7.0 m 3. (a) right; (b) more 4. (a) equal; (b) less
- Q** 1. c 3. b 5. (a) C'_1 ; (b) C'_1 7. (a) 4 s; (b) 3 s; (c) 5 s; (d) 4 s; (e) 10 s 9. (a) a tie of 3, 4, and 6, then a tie of 1, 2, and 5; (b) 1, then a tie of 2 and 3, then 4, then a tie of 5 and 6; (c) 1, 2, 3, 4, 5, 6; (d) 2 and 4; (e) 1, 2, 5 11. (a) 3, tie of 1 and 2, then 4; (b) 4, tie of 1 and 2, then 3; (c) 1, 4, 2, 3
- P** 1. 0.990 50 3. (a) 0.999 999 50 5. 0.446 ps 7. 2.68×10^3 y 9. (a) 87.4 m; (b) 394 ns 11. 1.32 m 13. (a) 26.26 y; (b) 52.26 y; (c) 3.705 y 15. (a) 0.999 999 15; (b) 30 ly 17. (a) 138 km; (b) -374 μs 19. (a) 25.8 μs ; (b) small flash 21. (a) $\gamma[1.00 \mu\text{s} - \beta(400 \text{ m})/(2.998 \times 10^8 \text{ m/s})]$; (d) 0.750; (e) $0 < \beta < 0.750$; (f) $0.750 < \beta < 1$; (g) no 23. (a) 1.25; (b) 0.800 μs 25. (a) 0.480; (b) negative; (c) big flash; (d) 4.39 μs 27. 0.81c 29. (a) 0.35; (b) 0.62 31. 1.2 μs 33. (a) 1.25 y; (b) 1.60 y; (c) 4.00 y 35. 22.9 MHz 37. 0.13c 39. (a) 550 nm; (b) yellow 41. (a) 196.695; (b) 0.999 987 43. (a) 1.0 keV; (b) 1.1 MeV 45. 110 km 47. 1.01×10^7 km 49. (a) 0.222 cm; (b) 701 ps; (c) 7.40 ps 51. 2.83mc 53. $\gamma(2\pi m/lqLB)$; (b) no; (c) 4.85 mm; (d) 15.9 mm; (e) 16.3 ps; (f) 0.334 ns 55. (a) 0.707; (b) 1.41; (c) 0.414 57. 18 smu/y 59. (a) 2.08 MeV; (b) -1.21 MeV 61. (d) 0.801 63. (a) $vt \sin \theta$; (b) $t[1 - (v/c) \cos \theta]$; (c) 3.24c 67. (b) +0.44c 69. (a) 1.93 m; (b) 6.00 m; (c) 13.6 ns; (d) 13.6 ns; (e) 0.379 m; (f) 30.5 m; (g) -101 ns; (h) no; (i) 2; (k) no; (l) both 71. (a) 5.4×10^4 km/h; (b) 6.3×10^{-10} 73. 189 MeV 75. 8.7×10^{-3} ly 77. 7 79. 2.46 MeV/c 81. 0.27c

83. (a) 5.71 GeV; (b) 6.65 GeV; (c) 6.58 GeV/c; (d) 3.11 MeV; (e) 3.62 MeV; (f) 3.59 MeV/c

85. 0.95c 87. (a) 256 kV; (b) 0.745c

CHAPTER 38

- CP** 1. b, a, d, c 2. (a) lithium, sodium, potassium, cesium; (b) all tie 3. (a) same; (b)-(d) x rays 4. (a) proton; (b) same; (c) proton 5. same
- Q** 1. (a) greater; (b) less 3. potassium 5. only e 7. none 9. (a) decreases by a factor of $(1/2)^{0.5}$; (b) decreases by a factor of 1/2 11. amplitude of reflected wave is less than that of incident wave 13. electron, neutron, alpha particle 15. all tie
- P** 1. (a) $2.1 \mu\text{m}$; (b) infrared 3. 1.0×10^{45} photons/s 5. 2.047 eV 7. 1.1×10^{-10} W 9. (a) 2.96×10^{20} photons/s; (b) 4.86×10^7 m; (c) 5.89×10^{18} photons/m² · s 11. (a) infrared; (b) 1.4×10^{21} photons/s 13. 4.7×10^{26} photons 15. 170 nm 17. 676 km/s 19. 1.3 V; (b) 6.8×10^2 km/s 21. (a) 3.1 keV; (b) 14 keV 23. (a) 2.00 eV; (b) 0; (c) 2.00 V; (d) 295 nm 25. (a) 382 nm; (b) 1.82 eV 27. (a) 2.73 pm; (b) 6.05 pm 29. (a) 8.57×10^{18} Hz; (b) 3.55×10^4 eV; (c) 35.4 keV/c 31. 300% 33. (a) $-8.1 \times 10^{-9}\%$; (b) $-4.9 \times 10^{-4}\%$; (c) -8.9%; (d) -66% 35. (a) 2.43 pm; (b) 1.32 fm; (c) 0.511 MeV; (d) 939 MeV 37. (a) 41.8 keV; (b) 8.2 keV 39. 44° 41. (a) 2.43 pm; (b) 4.11×10^{-6} ; (c) -8.67×10^{-6} eV; (d) 2.43 pm; (e) 9.78×10^{-2} ; (f) -4.45 keV 43. 7.75 pm 45. (a) 1.9×10^{-21} kg · m/s; (b) 346 fm 47. 4.3 μeV 49. (a) 1.24 μm ; (b) 1.22 nm; (c) 1.24 fm; (d) 1.24 fm 51. (a) 15 keV; (b) 120 keV 53. neutron 55. (a) 3.96×10^6 m/s; (b) 81.7 kV 63. 2.1×10^{-24} kg · m/s 67. (a) 9.02×10^{-6} ; (b) 3.0 MeV; (c) 3.0 MeV; (d) 7.33×10^{-8} ; (e) 3.0 MeV; (f) 3.0 MeV 69. (a) -20%; (b) -10%; (c) +15% 71. 5.9 μeV 75. (a) 73 pm; (b) 3.4 nm; (c) yes, their average de Broglie wavelength is smaller than their average separation 79. 1.7×10^{-35} m 81. 0.19 m

CHAPTER 39

- CP** 1. b, a, c 2. (a) all tie; (b) a, b, c 3. a, b, c, d 4. $E_{1,1}$ (neither n_x nor n_y can be zero) 5. (a) 5; (b) 7
- Q** 1. a, c, b 3. (a) 18; (b) 17 5. equal 7. c 9. (a) decrease; (b) increase 11. $n = 1, n = 2, n = 3$ 13. (a) $n = 3$; (b) $n = 1$; (c) $n = 5$ 15. b, c, and d
- P** 1. 1.41 3. 0.65 eV 5. 0.85 nm 7. 1.9 GeV 9. (a) 72.2 eV; (b) 13.7 nm; (c) 17.2 nm; (d) 68.7 nm; (e) 41.2 nm; (g) 68.7 nm; (h) 25.8 nm 11. (a) 13; (b) 12 13. (a) 0.020; (b) 20 15. (a) 0.050; (b) 0.10; (c) 0.0095 17. 56 eV 19. 109 eV 23. 3.21 eV 25. 1.4×10^{-3} 27. (a) 8; (b) 0.75; (c) 1.00; (d) 1.25; (e) 3.75; (f) 3.00; (g) 2.25 29. (a) 7; (b) 1.00; (c) 2.00; (d) 3.00; (e) 9.00; (f) 8.00; (g) 6.00 31. 4.0 33. (a) 12.1 eV; (b) 6.45×10^{-27} kg · m/s; (c) 102 nm 35. (a) 291 nm^{-3} ; (b) 10.2 nm^{-1} 41. (a) 0.0037; (b) 0.0054 43. (a) 13.6 eV; (b) -27.2 eV 45. (a) $(r^4/8a^5)[\exp(-r/a)]\cos^2 \theta$; (b) $(r^4/16a^5)[\exp(-r/a)]\sin^2 \theta$ 47. 4.3×10^3 49. (a) 13.6 eV; (b) 3.40 eV 51. 0.68 59. (b) $(2\pi/h)[2m(U_0 - E)]^{0.5}$ 61. (b) meter^{-2.5} 63. (a) n ; (b) $2\ell + 1$; (c) n^2

CHAPTER 40

- CP** 1. 7 2. (a) decrease; (b)-(c) remain the same 3. A, C, B
- Q** 1. (a) 2; (b) 8; (c) 5; (d) 50 3. all true 5. same number (10) 7. 2, -1, 0, and 1 9. (a) 2; (b) 3 11. (a) n ; (b) n and ℓ 13. In addition to the quantized energy, a helium atom has kinetic energy; its total energy can equal 20.66 eV.
- P** 1. 24.1° 3. (a) 3.65×10^{-34} J · s; (b) 3.16×10^{-34} J · s 5. (a) 3; (b) 3 7. (a) 4; (b) 5; (c) 2 9. (a) 3.46; (b) 3.46; (c) 3; (d) 3; (e) -3; (f) 30.0°; (g) 54.7°; (h) 150° 13. 72 km/s² 15. (a) 54.7°; (b) 125° 17. 19 mT 19. 5.35 cm 21. 44 23. 42 25. (a) 51; (b) 53; (c) 56

AN-12 CHAPTER 9 CENTER OF MASS AND LINEAR MOMENTUM

- 27.** (a) $(2, 0, 0, +\frac{1}{2})$, $(2, 0, 0, -\frac{1}{2})$; (b) $(2, 1, 1, +\frac{1}{2})$, $(2, 1, 1, -\frac{1}{2})$,
 $(2, 1, 0, +\frac{1}{2})$, $(2, 1, 0, -\frac{1}{2})$, $(2, 1, -1, +\frac{1}{2})$, $(2, 1, -1, -\frac{1}{2})$ **29.** g
31. (a) $4p$; (b) 4; (c) $4p$; (d) 5; (e) $4p$; (f) 6 **33.** 12.4 kV **35.** (a) 35.4 pm;
(b) 56.5 pm; (c) 49.6 pm **39.** 0.563 **41.** 80.3 pm **43.** (a) 69.5 kV;
(b) 17.8 pm; (c) 21.3 pm; (d) 18.5 pm **45.** (a) 49.6 pm; (b) 99.2 pm
47. $2.0 \times 10^{16} \text{ s}^{-1}$ **49.** 2×10^7 **51.** 9.0×10^{-7} **53.** $7.3 \times 10^{15} \text{ s}^{-1}$
55. (a) 3.60 mm; (b) 5.24×10^{17} **57.** (a) 0; (b) 68 J **59.** 3.0 eV
61. (a) 3.03×10^5 ; (b) 1.43 GHz; (d) 3.31×10^{-6} **63.** 186
65. (a) 2.13 meV; (b) 18 T **69.** (a) no; (b) 140 nm **71.** $n > 3$;
 $\ell = 3$; $m_\ell = +3, +2, +1, 0, -1, -2, -3$; $m_s = \pm \frac{1}{2}$ **73.** (a) 6.0;
(b) $3.2 \times 10^6 \text{ y}$ **75.** argon

CHAPTER 41

CP 1. larger **2.** a, b, and c

- Q** 1. b, c, d (the latter due to thermal expansion) **3.** 8
5. below **7.** increase **9.** much less than **11.** b and d
P 3. $8.49 \times 10^{28} \text{ m}^{-3}$ **5.** (b) $6.81 \times 10^{27} \text{ m}^{-3} \text{ eV}^{-3/2}$;
(c) $1.52 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$ **7.** (a) 0; (b) 0.0955 **9.** (a) $5.86 \times 10^{28} \text{ m}^{-3}$;
(b) 5.49 eV; (c) $1.39 \times 10^3 \text{ km/s}$; (d) 0.522 nm **11.** (a) $1.36 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$;
(b) $1.68 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$; (c) $9.01 \times 10^{27} \text{ m}^{-3} \text{ eV}^{-1}$;
(d) $9.56 \times 10^{26} \text{ m}^{-3} \text{ eV}^{-1}$; (e) $1.71 \times 10^{18} \text{ m}^{-3} \text{ eV}^{-1}$ **13.** (a) 6.81 eV;
(b) $1.77 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$; (c) $1.59 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$ **15.** (a) $2.50 \times 10^3 \text{ K}$;
(b) $5.30 \times 10^3 \text{ K}$; **17.** 3 **19.** (a) 1.0; (b) 0.99; (c) 0.50; (d) 0.014;
(e) 2.4×10^{-17} ; (f) $7.0 \times 10^2 \text{ K}$ **21.** (a) 0.0055; (b) 0.018
25. (a) 19.7 kJ; (b) 197 s **27.** (a) $1.31 \times 10^{29} \text{ m}^{-3}$; (b) 9.43 eV;
(c) $1.82 \times 10^3 \text{ km/s}$; (d) 0.40 nm **29.** 57.1 kJ **31.** (a) 226 nm;
(b) ultraviolet **33.** (a) 1.5×10^{-6} ; (b) 1.5×10^{-6} **35.** 0.22 μg
37. (a) 4.79×10^{-10} ; (b) 0.0140; (c) 0.824 **39.** 6.0 $\times 10^5$ **41.** 4.20 eV
43. $13 \mu\text{m}$ **47.** (a) 109.5° ; (b) 238 pm **49.** (b) $1.8 \times 10^{28} \text{ m}^{-3} \text{ eV}^{-1}$
53. $3.49 \times 10^3 \text{ atm}$

CHAPTER 42

- CP** 1. ${}^{90}\text{As}$ and ${}^{158}\text{Nd}$ **2.** a little more than 75 Bq (elapsed time is a little less than three half-lives) **3.** ${}^{206}\text{Pb}$
Q 1. (a) ${}^{196}\text{Pt}$; (b) no **3.** yes **5.** (a) less; (b) greater **7.** ${}^{240}\text{U}$
9. no effect **11.** yes **13.** (a) all except ${}^{198}\text{Au}$; (b) ${}^{132}\text{Sn}$ and ${}^{208}\text{Pb}$
15. d
P 1. $1.3 \times 10^{-13} \text{ m}$ **3.** 46.6 fm **5.** (a) 0.390 MeV; (b) 4.61 MeV
7. (a) $2.3 \times 10^{17} \text{ kg/m}^3$; (b) $2.3 \times 10^{17} \text{ kg/m}^3$; (d) $1.0 \times 10^{25} \text{ C/m}^3$;
(e) $8.8 \times 10^{24} \text{ C/m}^3$ **9.** (a) 6; (b) 8 **11.** (a) 6.2 fm; (b) yes
13. 13 km **17.** 1.0087 u **19.** (a) 9.303%; (b) 11.71%
21. (b) 7.92 MeV/nucleon **25.** 5.3×10^{22} **27.** (a) 0.250; (b) 0.125
29. (a) 64.2 h; (b) 0.125; (c) 0.0749 **31.** (a) $7.5 \times 10^{16} \text{ s}^{-1}$;
(b) $4.9 \times 10^{16} \text{ s}^{-1}$ **33.** $1 \times 10^{13} \text{ atoms}$ **37.** 265 mg

- 39.** (a) $8.88 \times 10^{10} \text{ s}^{-1}$; (b) 1.19×10^{15} ; (c) $0.111 \mu\text{g}$ **41.** $1.12 \times 10^{11} \text{ y}$
43. $9.0 \times 10^8 \text{ Bq}$ **45.** (a) $3.2 \times 10^{12} \text{ Bq}$; (b) 86 Ci **47.** (a) 2.0×10^{20} ;
(b) $2.8 \times 10^9 \text{ s}^{-1}$ **49.** (a) 1.2×10^{-17} ; (b) 0 **51.** 4.269 MeV
53. 1.21 MeV **55.** 0.783 MeV **57.** (b) 0.961 MeV **59.** 78.3 eV
61. (a) 1.06×10^{19} ; (b) 0.624×10^{19} ; (c) 1.68×10^{19} ; (d) $2.97 \times 10^9 \text{ y}$
63. 1.7 mg **65.** 1.02 mg **67.** 2.50 mSv **69.** (a) 6.3×10^{18} ;
(b) 2.5×10^{11} ; (c) 0.20 J; (d) 2.3 mGy; (e) 30 mSv
71. (a) 6.6 MeV; (b) no **73.** (a) 25.4 MeV; (b) 12.8 MeV;
(c) 25.0 MeV **75.** ${}^7\text{Li}$ **77.** $3.2 \times 10^4 \text{ y}$ **79.** 730 cm²
81. ${}^{225}\text{Ac}$ **83.** 30 MeV **89.** 27 **91.** (a) 11.906 83 u; (b) 236.2025 u
93. 600 keV **95.** (a) 59.5 d; (b) 1.18

CHAPTER 43

- CP** 1. c and d **2.** (a) no; (b) yes; (c) no **3.** e
Q 1. (a) 101; (b) 42 **3.** ${}^{239}\text{Np}$ **5.** ${}^{140}\text{I}$, ${}^{105}\text{Mo}$, ${}^{152}\text{Nd}$, ${}^{123}\text{In}$, ${}^{115}\text{Pd}$
7. increased **9.** less than **11.** still equal to 1
P 1. (a) 16 day⁻¹; (b) 4.3×10^8 **3.** 4.8 MeV **5.** $1.3 \times 10^3 \text{ kg}$
 $7.3.1 \times 10^{10} \text{ s}^{-1}$ **9.** (a) 2.6×10^{24} ; (b) $8.2 \times 10^{13} \text{ J}$; (c) $2.6 \times 10^4 \text{ y}$
11. -23.0 MeV **13.** (a) 251 MeV; (b) typical fission energy is
200 MeV **15.** (a) 84 kg; (b) 1.7×10^{25} ; (c) 1.3×10^{25} **17.** (a) ${}^{153}\text{Nd}$;
(b) 110 MeV; (c) 60 MeV; (d) $1.6 \times 10^7 \text{ m/s}$; (e) $8.7 \times 10^6 \text{ m/s}$
21. 557 W **23.** 0.99938 **25.** (b) 1.0; (c) 0.89; (d) 0.28; (e) 0.019;
(f) 8 **27.** (a) 75 kW; (b) $5.8 \times 10^3 \text{ kg}$ **29.** $1.7 \times 10^9 \text{ y}$ **31.** 170 keV
33. 1.41 MeV **35.** 10^{-12} m **37.** (a) $4.3 \times 10^9 \text{ kg/s}$; (b) 3.1×10^{-4}
41. $1.6 \times 10^8 \text{ y}$ **43.** (a) 24.9 MeV; (b) 8.65 megatons TNT
45. (a) $1.8 \times 10^{38} \text{ s}^{-1}$; (b) $8.2 \times 10^{28} \text{ s}^{-1}$ **47.** (a) 4.1 eV/atom;
(b) 9.0 MJ/kg; (c) $1.5 \times 10^3 \text{ y}$ **49.** 14.4 kW **51.** ${}^{238}\text{U} + n \rightarrow {}^{239}\text{U} \rightarrow$
 ${}^{239}\text{Np} + e + \nu$; ${}^{239}\text{Np} \rightarrow {}^{239}\text{Pu} + e + \nu$ **55.** (a) $3.1 \times 10^{31} \text{ protons/m}^3$;
(b) 1.2×10^6

CHAPTER 44

- CP** 1. (a) the muon family; (b) a particle; (c) $L_\mu = +1$
2. b and e **3.** c
Q 1. b, c, d **3.** (a) 1; (b) positively charged **5.** a, b, c, d **7.** d
9. c **11.** (a) lepton; (b) antiparticle; (c) fermion; (d) yes
P 1. $\pi^- \rightarrow \mu^- + \bar{\nu}$ **3.** 2.4 pm **5.** 2.4×10^{-43} **7.** 7.769 MeV
9. 2.7 cm/s **11.** (a) angular momentum, L_c ; (b) charge, L_μ ;
(c) energy, L_μ **15.** (a) energy; (b) strangeness; (c) charge
17. (a) yes; (b) –(d) no **19.** (a) 0; (b) -1 ; (c) 0 **21.** (a) K^+ ; (b) \bar{n} ; (c)
 K^0 **23.** (a) not possible; (b) uuu **25.** (a) $u\bar{u}d\bar{d}$; (b) $\bar{u}\bar{d}\bar{d}$
27. $s\bar{d}$ **29.** (a) Ξ^0 ; (b) Σ^- **31.** $2.77 \times 10^8 \text{ ly}$ **33.** 668 nm
35. $1.4 \times 10^{10} \text{ ly}$ **37.** (b) 0.934; (c) $1.28 \times 10^{10} \text{ ly}$ **39.** (a) 0.26 meV;
(b) 4.8 mm **41.** (a) 121 m/s; (b) 0.00406; (c) 248 y **43.** (a) 2.6 K; (b)
976 nm **45.** $1.08 \times 10^{42} \text{ J}$ **49.** (a) 0.785c; (b) 0.993c;
(c) C2; (d) Cl; (e) 51 ns; (f) 40 ns

Figures are noted by page numbers in *italics*, tables are indicated by *t* following the page number.

A

- a_g (gravitational acceleration), 334
 - variation with altitude, 334t
- absolute pressure, 363
- absolute zero, 476
- absorption, of heat by solids and liquids, 485–488
- absorption event, photons, 1058
 - in lasers, 1132
- absorption lines, 1100, 1101
- ac (alternating current), 826, 835
 - acceleration, 263t. *See also* angular acceleration; forces
 - constant, 22, 22–24
 - free-fall, 25–26
 - graphical integration in motion analysis, 27, 27–28
 - negative, 20
 - and Newton's first law, 87–91
 - Newton's laws applications, 100–105
 - and Newton's second law, 91–94
 - one-dimensional motion, 18–26
 - principle of equivalence (with gravitation), 347
 - projectile motion, 64–70
 - relating linear to angular, 251, 251
 - relative motion in one dimension, 73–74
 - relative motion in two dimensions, 74–75
 - rockets, 225, 224–226
 - rolling down ramp, 278–279, 279
 - sign of, 20
 - simple harmonic motion, 389, 389
 - system of particles, 206–209
 - two- and three-dimensional motion, 73–75
 - uniform circular motion, 70–72, 71, 124
 - as vector quantity, 38
 - yo-yo, 281
 - acceleration amplitude, in simple harmonic motion, 389
 - acceleration vectors, 38
 - accelerators, 747–749, 1219, 1220, 1220
 - acceptor atoms, 1153

acre-foot, 9

- action at a distance, 580
- activity, of radioactive sample, 1175
- addition
 - of vectors by components, 44–46
 - of vectors geometrically, 39, 39–40, 40
- adiabat, 526, 526
- adiabatic expansion, 491–492, 492
 - ideal gas, 526, 526–528
- adiabatic processes
 - first law of thermodynamics for, 491–493, 492t
 - summarized, 528, 528t
- adiabatic wind, 534
- air
 - bulk modulus, 446
 - density, 360t
 - dielectric properties at 1 atm, 669, 669t
 - and drag force, 121–123
 - effect on projectile motion, 68–69, 68
 - electric breakdown, 593, 593, 602
 - electric breakdown field, 580t
 - index of refraction at STP, 906t
 - speed of sound in, 446–447, 447t
 - terminal speeds in, 122t
 - thermal conductivity, 494t
- air conditioners, 548
- airspeed, 84
- alpha decay, 1177–1179, 1178
- alpha particles, 572, 644, 1166–1168, 1166, 1177
- binding energy per nucleon, 1172
- magic nucleon number, 1186
- radiation dosage, 1184
- in thermonuclear fusion, 1209
- alternating current (ac), 826, 835
 - See also* transformers
- alternating current circuits, 835. *See also* transformers
 - capacitive load, 838–840, 839
 - forced oscillations, 836, 836
 - inductive load, 840–841, 841
 - LC oscillations, 826–829, 827
 - leading and lagging in, 841, 842t, 843

phase and amplitude relationships, 842t

- power in, 847–849
- resistive load, 836–838, 837
- RLC circuit damped oscillations, 833, 833–834
- series RLC circuits, 842–847, 843
- alternating current generator, 835
 - with capacitive load, 838–840, 839
 - with inductive load, 840, 840–842, 841
 - with resistive load, 836–838, 837
 - with transformer, 850, 850–853
- ammeters, 720, 720
- Ampère, André-Marie, 771
- Ampere-Maxwell law, 864, 864–866
 - Maxwell's equation form, 869t
- Ampere's law, 771–774
- ampere (unit), 565, 683, 770
- Amperian loop, 771, 771–774
- amplitude
 - alternating current, 842t
 - of emf in ac, 835
 - exponentially decaying in RLC circuits, 834
 - LC oscillations, 826
 - simple harmonic motion, 387–388, 388
 - waves, 415, 415–416, 416
- amplitude ratio, traveling electromagnetic waves, 892
- amusement park rides
 - roller coaster headache, 252
 - roller coasters, 20
 - Rotor, 249
- analyzer, 903
- Andromeda Galaxy, 330–331, 331
- angle of incidence, 906, 906
- angle of minimum deviation
 - prism, 920
 - rainbow, 922
- angle of reflection, 906, 906
- angle of refraction, 906, 906
- angles, 43
- angular acceleration, 243, 263t
 - relating to linear, 251, 251
 - rolling wheel, 278, 278
 - rotation with constant, 248–249
- angular amplitude
- simple pendulum, 396
- angular displacement, 242, 242
- nonvector nature of, 247
- angular frequency
 - circulating charged particle, 745
 - damped harmonic oscillator, 400
 - of emf, 835
 - LC oscillations, 826, 831–832, 832
 - natural, 402
 - simple harmonic motion, 386–389, 388
 - sound waves, 449
 - waves, 416
- angular magnification
 - compound microscope, 944
 - refracting telescope, 945
 - simple magnifying lens, 943
- angular momentum, 284, 284–285, 289t
 - conservation of, 290–293, 291, 292
 - at equilibrium, 305
 - Newton's second law in angular form, 285
 - rigid body rotating about fixed axis, 288–290, 289
 - system of particles, 288
- angular motion, 242
- angular position, 242–243, 242, 263t
 - relating to linear, 250
- angular simple harmonic motion, 394, 394–395
- angular simple harmonic oscillator, 394, 394–395
- angular speed, 243, 244
 - relating to linear, 250–251
 - in rolling, 275–277, 276
- angular velocity, 243–246, 263t
 - vector nature of, 247, 246–247
- angular wave number, 415, 1072
 - sound waves, 449
- annihilation, 1222
 - electron-positron, 572, 572, 1222
 - proton-antiproton, 1224t, 1224
- annular cylinder
 - rotational inertia, 255t
- antenna, 891, 891
- antibaryons, 1237
- antiderivative, 24
- antihydrogen, 1108, 1222

antimatter, 1222
 energy released by 1 kg, 1195
 antineutrino, 1180
 antinodes, 431, 431, 433
 antiparticles, 1222, 1224
 formation in early universe, 1240
 antiprotons, 1222
 antihydrogen, 1108, 1222
 discovery of, 1248
 proton-antiproton annihilation, 1224t, 1224
 antisolar point, 909, 909
 aphelion distance, 344
 apparent weight, 96
 in fluids, 369–370
 Archimedes' principle, 367, 367–370
 area vector, 606
 astronomical Doppler effect, 1040–1042, 1041
 astronomical unit, 12
 atmosphere (atm), 361
 atmospheric pressure, 361t
 atomic bomb, 1172, 1201
 rating, 1215
 as trigger for fusion bomb, 1211
 atomic clocks, 5–6
 time dilation tests, 1029
 atomic mass, 1171
 selected nuclides, 1168t
 atomic mass units, 7, 1171
 atomic number, 572, 1117, 1168
 atoms, 1083. *See also* electrons; neutrons; protons
 absorption and emission of light, 1113
 angular momentum, 1113–1114, 1113, 1115–1117, 1116
 Bohr model, 1096–1099, 1097
 Einstein-de Haas experiment, 1114, 1114
 electron spin, 1115
 formation in early universe, 1240
 magnetic resonance, 1121, 1120–1121
 magnetism, 1113–1114, 1113
 matter wave interference, 1068, 1068
 orbital angular momentum, 1115–1117, 1116, 1117
 properties of, 1112–1114, 1113
 spin angular momentum, 1116–1117, 1117

Stern-Gerlach experiment, 1118, 1118–1120
 attractive forces, 331
 aurora, 561
 automobiles. *See also* race cars
 magnet applications, 735
 spark discharge from, 645, 645
 tire pressure, 361t
 autotransformer, 859, 859
 average acceleration
 one-dimensional motion, 18
 two- and three-dimensional motion, 62–64
 average angular acceleration, 243
 average angular velocity, 243
 average power, 155, 185
 engines, 547
 traveling wave on stretched string, 423
 average speed
 of gas molecules, 516–518
 one-dimensional motion, 16
 average velocity
 constant acceleration, 23
 one-dimensional motion, 14–15
 two- and three-dimensional motion, 60
 Avogadro's number, 507–508, 685
 axis of rotation, 242, 242

B

Babinet's principle, 1016
 ball
 projectile motion, 64–66, 65
 ballet dancing
 grand jeté, 208, 208
 tour jeté, 292, 292
 ballistic pendulum, 220, 220
 balloons, lifting capacity, 534
 Balmer series, 1097, 1099–1101, 1100
 band-gap pattern
 crystalline solid, 1144
 insulator, 1144
 metal, 1145
 semiconductor, 1151
 bands, energy bands in crystalline solids, 1144, 1144
 bar magnets
 Earth as, 870, 870
 magnetic dipole moment of small, 755, 755t
 magnetic field, 862, 862–863
 magnetic field lines, 738, 738–739
 barrier tunneling, 1074, 1074–1076

baryonic matter, 1240
 baryon number, conservation of, 1229
 baryons, 1221, 1228–1229, 1245
 conservation of baryon number, 1229
 and eightfold way, 1231t, 1231–1232
 and quark model, 1232, 1237
 baseball
 collision of ball with bat, 211, 211
 fly ball, air resistance to, 68, 68
 base quantities, 1
 base standards, 1
 basilisk lizards, 233, 233
 basketball free throws, 58
 bats, navigation using ultrasonic waves, 464
 batteries. *See also* emf
 connected to capacitors, 658, 658, 666
 and current, 682–684, 683
 as emf devices, 705, 706–707
 in multiloop circuits, 714, 714–719
 multiple batteries in multiloop circuit, 716–717, 717
 potential difference across, 711–712, 712
 and power in circuits, 695, 695–696
 in RC circuits, 720, 720–724
 recharging, 712
 in RL circuits, 808–810
 in single-loop circuits, 707–710
 work and energy, 706, 706–707
 beam, 892
 beam expander, 957
 beam splitter, 980, 1067, 1067
 beats, 459, 459–460
 becquerel, 1175
 bends, 507
 during flight, 379
 Bernoulli's equation, 374–377
 beta decay, 577, 1180–1182, 1181
 and quark model, 1234
 beta-minus decay, 1180
 beta-plus decay, 1180
 bi-concave lens, 957
 bi-convex lens, 957
 bicycle wheels
 rolling, 275–277, 276, 277
 rolling with friction, 278–280, 279
 Big Bang, 1237, 1240–1242, 1241
 binding energy, *See* nuclear binding energy
 binding energy per nucleon, 1171, 1172, 1198
 Biot-Savart law, 765–766, 771, 779
 bivalent atom, 1145
 blackbody radiator, 496
 black holes, 330
 event horizon, 337
 gravitational lensing caused by, 348, 348
 miniature, 351
 and quasars, 1239
 supermassive, 330
 blocks
 acceleration of falling, 261
 connected to massless, frictionless pulleys, 97, 98, 100, 100–101
 floating, 370
 forces on stationary, 117, 116–118
 friction of sliding, 97, 97
 hanging and sliding, 100, 100–101
 Newton's laws applications, 92, 100–105
 normal forces, 96–97, 97
 power used in work on, 156, 156
 stable static equilibrium, 305–306, 307, 309–313
 third-law force pair, 98–99, 99
 work done by external force with friction, 181, 181
 block-spring oscillator, 830
 block-spring systems
 damped oscillating systems, 400, 400–401
 and electrical-mechanical analogy, 829t, 829
 kinetic energy, 149, 149–151, 151
 oscillating systems, 390, 390–392
 potential energy, 167–168, 167, 171
 blood pressure
 normal systolic, 360t
 blue shift, 1041
 bob, of pendulum, 395
 body armor, 444, 444
 body diagonal, 56
 Bohr, Niels, 1089, 1185, 1199
 Bohr magneton, 872, 873, 1117
 Bohr model, of hydrogen, 1096–1199, 1097

- Bohr radius, 1098, 1106
 boiling point, 486
 selected substances, 487t
 Boltzmann, Ludwig, 1133
 entropy equation, 553
 Boltzmann constant, 508
 Bose, Satyendra Nath, 1220
 Bose-Einstein condensate, 1221, 1221
 bosons, 1220–1221, 1221
 bottomness, 1231
 bottom quark, 1233t, 1234
 boundary condition, 1103
 Bragg angle, 1012–1013
 Bragg's law, 1012–1013
 Brahe, Tycho, 342
 branches, circuits, 714
 breakdown potential, 670
 breakeven, in magnetic confinement, 1212
 Brewster angle, 912, 912
 Brewster's law, 913
 bright bands, *See* bright fringes
 bright fringes
 double-slit interference, 964, 964, 966
 single-slit diffraction, 992–993, 993
 British thermal unit (Btu), 485
 Brookhaven accelerator, 1219
 bubble chambers, 572, 572–573, 737, 737
 gamma ray track, 1069, 1069–1070
 proton-antiproton annihilation event, 1223, 1223
 buildings
 natural angular frequency, 402
 submerged during earthquakes, 7–8
 swaying in wind, 393–394, 434
 bulk modulus, 317, 446
 bungee-cord jumping, 166, 166
 buoyant force, 367–370, 368
- C**
c. *See* speed of light
 calorie (cal), 484
 Calorie (Cal) (nutritional), 484
 cameras, 943, 955
 cancer radiation therapy, 1165
 with ^{60}Co , 1190
 capacitance, 656–658
 calculating, 659–662
 LC oscillations, 826–829, 827
 parallel circuits, 662–663, 715t
 RC circuits, 720, 720–724
 RLC circuits, 833–834, 833
 cascade, of decay processes, 1230
 cat, terminal speed of falling, 122, 122–123
 cathode ray tube, 740, 740–741
 Celsius temperature scale, 479–481, 480
 center of curvature
- series circuits, 663–666, 715t, 842–844
 capacitive reactance, 838
 capacitive time constant
 RC circuits, 721, 721
 capacitors, 656, 656–658, 658
 with ac generator, 838, 838–840
 charging, 657, 658
 charging in *RC* circuits, 720, 720–721
 with dielectrics, 669, 669–671
 discharging, 658
 discharging in *RC* circuits, 720, 720–721, 722–724
 displacement current, 866–869, 867
 electric field calculation, 659
 energy density, 668
 energy stored in electric field, 667–669
 Faraday's, 669, 669
 induced magnetic field, 863, 865–866
 LC oscillations, 826–829, 827
 in parallel, 662–663, 663, 715t
 phase and amplitude relationships for ac circuits, 842t
 potential difference calculation, 659–662
 RC circuits, 720, 720–724
 RLC circuits, 833–834, 833
 in series, 663–667, 664, 715t, 842–844, 843
 series *RLC* circuits, 842–844, 843
 carbon cycle, in stars, 1217
 carbon¹⁴ dating, 1183
 carbon dioxide
 molar specific heat at constant volume, 520t
 RMS speed at room temperature, 512t
 carbon disulfide
 index of refraction, 906t
 Carnot cycle, 543, 543, 544
 Carnot engines, 543–546, 543
 efficiency, 545–546, 549–550
 Carnot refrigerators, 548, 549–550
 cars, *See* automobiles; race cars
 cascade, of decay processes, 1230
 cat, terminal speed of falling, 122, 122–123
 cathode ray tube, 740, 740–741
 Celsius temperature scale, 479–481, 480
 center of curvature
- spherical mirrors, 928, 928
 spherical refracting surfaces, 934, 934
 center of gravity, 308–309, 308
 center of mass
 and center of gravity, 308–309
 one-dimensional inelastic collisions, 218–220, 219, 220
 rolling wheel, 276, 276
 solid bodies, 203–206, 205
 system of particles, 202–203, 202, 206–209
 center of momentum frame, 1056
 center of oscillation, physical pendulum, 397
 centigrade temperature scale, 479
 central axis
 spherical mirrors, 929, 929
 central configuration peak, 552
 central diffraction maximum, 998, 998
 central interference maximum, 966
 central line, 1007
 central maximum
 diffraction patterns, 990, 990
 centripetal acceleration, 70–71
 centripetal force, 124–129, 125
 Cerenkov counters, 1248
 CERN accelerator, 1219, 1236
 antihydrogen, 1222
 pion beam experiments, 1024
 cesium
 work function, 1079
 cesium₁₃₇, in fallout, 1194
 chain-link conversion, of units, 3
 chain reaction, 1201
 change in entropy, *See* entropy changes
 characteristic nuclear time, 1194
 characteristic x-ray spectrum, 1128, 1128
 charge, *See* electric charge
 charge carrier density, 685
 silicon and copper compared, 696t, 696–697
 charge carriers, 684, 1151
 doped semiconductors, 1152–1154
 drift speed, 686, 685–688
 semiconductors, 696
 charged disk
 electric field due to, 591
- electric potential due to, 640, 640
 charged isolated conductor
 with cavity, 612, 613
 electric potential, 644–645, 645
 in external electric field, 645, 645
 Gauss' law for, 612–615
 spark discharge from, 561–562
 charged objects, 561–562, 580
 charge number, 1117
 charge quantum number, 1224
 charging
 capacitors, 657–658
 capacitors in *RC* circuits, 720, 720–721
 charm, 1231
 charm quark, 1233t, 1234
 Chernobyl reactor, 1194
 chip (integrated circuits), 1159–1160
 chocolate crumb explosion mystery, 626, 653, 680, 703
 chromatic aberration, 946
 chromatic dispersion, 907–908, 908
 circuit elements, 658
 circuits, 658, 658, 705, 715t. *See also* alternating current circuits
 capacitive load, 838–840, 839
 capacitors in parallel, 662–663, 663, 715t
 capacitors in series, 663, 663–666, 715t, 842–844, 843
 current direction, 714
 current in single-loop, 707, 707–709
 direct-current (DC), 705
 grounding, 712, 712
 inductive load, 840–842, 841
 inductors in series, 842–844, 843
 integrated, 1159–1160, 1159
 internal resistance, 709, 709
 multiloop, 708, 714, 714–719
 oscillating, 826
 parallel capacitors, 662–663, 663, 715t
 parallel inductors, 822
 parallel resistors, 714, 714–719, 715t
 potential difference between two points, 711, 711–714, 712
 power, potential, and emf, 712–713

circuits (*cont.*)
 power in, 695–696
RC, 720, 720–724
 resistive load, 836–838, 837
 resistors in parallel, 714,
 714–719, 715t
 resistors in series, 710, 710,
 715t, 842–844, 843
RL, 807, 807–810, 809
RLC, 833, 833–834
 series capacitors, 663,
 663–666, 715t, 842–844,
 843
 series inductors, 842–844,
 843
 series resistors, 710, 710,
 715t, 842–844, 843
 single-loop, 707–710, 836
 work, energy, and emf, 706,
 706–707

circular aperture
 diffraction patterns, 1000,
 1000–1003

circular arc charge distributions, 590

clocks
 event measurement array,
 1025, 1025
 time dilation tests,
 1029–1030, 1057

closed circuit, 709, 709

closed cycle processes
 first law of thermodynamics
 for, 492, 492t

closed path, 168–169, 169

closed-path test, for conservative force, 168

closed shell, 1186

closed subshell, 1126

closed system, 207
 entropy, 542
 linear momentum conservation, 215

COBE (Cosmic Background Explorer) satellite, 1242

coefficient of kinetic friction, 119–121

coefficient of linear expansion, 482
 selected materials, 482t

coefficient of performance refrigerators, 548

coefficient of static friction, 119–121

coefficient of volume expansion, 482

coherence, 969

coherence length, 1131

coherent light, 969, 1131

coils, 753. *See also* inductors

in ideal transformers, 850,
 850–851
 magnetic field, 779, 779–780
 mutual induction, 813–815,
 814
 self-induction, 806, 806–807

cold-weld, 118, 118

collective model, of nucleus, 1185

collimated slit, 1118

collimator, 1008, 1118

collisions, 211–214
 elastic in one dimension,
 221, 221–223
 glancing, 224, 224
 impulse of series of collisions, 213–214, 214
 impulse of single collision, 212, 211–212
 inelastic in one dimension, 218, 218–220, 219
 momentum and kinetic energy, 217–218
 two dimensions, 224, 224

color force, 1236–1237

color-neutral quarks, 1236–1237

color-shifting inks, 958

comets
 dust and ion tails, 917, 917
 stimulated emission from, 1140–1141, 1141

commutator, 752, 752

compass, 869

completely inelastic collisions, 218, 218–220, 219

complex conjugate, 1071

component notation, vectors, 42

components, of light, 907–908

components, of vectors, 41, 41–44
 adding vectors by, 45, 44–46

composite slab, conduction through, 494–495, 495

compound microscope, 944, 944

compound nucleus, 1185

compressibility, 318, 361

compressive stress, 316–317

Compton scattering, 1062, 1062–1065, 1063

Compton shift, 1062, 1062–1065

Compton wavelength, 1064

computers
 preventing sparks when working on, 732

concave lenses, 957

concave mirrors, 929, 929
 images, 931

concrete
 coefficient of linear expansion, 482t
 elastic properties, 317, 317t

condensing, 486

conducting devices, 692

conducting plates
 eddy currents, 800
 Gauss' law, 617, 617–618

conduction, 494–495, 494, 495
 conduction band, 1151, 1151

conduction electrons, 563, 682, 1145–1150

conduction rate, 494

conductivity, 690, 1146

conductors, 563–564, 682. *See also* electric current
 drift speed in, 685–688
 Ohm's law, 692–695

configurations, in statistical mechanics, 550–553

confinement principle, 1084

conservation of angular momentum, 290–294, 291, 292

conservation of baryon number, 1229

conservation of electric charge, 572–573

conservation of energy, 140, 183–186, 184

conservation of lepton number, 1228

conservation of linear momentum, 215–217

conservation of strangeness, 1231

conservative forces, 168, 629
 path independence of, 168–170, 169

constant acceleration
 one-dimensional motion, 22, 22–24

constant angular acceleration
 rotation with, 248–249

constant linear acceleration, 248

constant-pressure molar specific heat, 521–522

constant-pressure processes, 489, 489–490
 summarized, 528, 528t

work done by ideal gases, 511

constant-pressure specific heat, 486

constant-temperature processes
 summarized, 528, 528t

work done by ideal gases, 509, 509–510

constant-volume gas thermometer, 478, 478–479

constant-volume molar specific heat, 519–520

constant-volume processes, 489, 489–490
 first law of thermodynamics for, 492t, 492
 summarized, 528, 528t

work done by ideal gases, 510

constant-volume specific heat, 486

contact potential difference, 1155

continuous charge distribution, 586
 electric potential due to, 639, 639–640, 640

continuous x-ray spectrum, 1127, 1127–1128

contracted length, 1032

convection, 495

converging lens, 936–937, 937

conversion factors, 3

convex lenses, 957

convex mirrors, 929, 929
 images, 931

Coordinated Universal Time (UTC), 5–6

copper
 coefficient of linear expansion, 482t
 conduction electrons, 563
 electric properties compared to silicon, 696t, 696–697, 1143t, 1150–1151
 energy levels, 1143, 1143
 Fermi energy, 1145, 1162
 Fermi speed, 1145
 heats of transformation, 487t
 mean free time, 694
 resistivity, 689t, 691, 691, 1151
 rubbing rod with wool, 563

temperature coefficient of resistivity, 1151

unit cell, 1142, 1142

copper wire
 aluminum competes with, 700
 as conductor, 563, 563, 683, 683
 drift speed in, 686
 electric field inside, 580t
 magnetic force on current carrying, 750, 750–751

cord (unit of wood), 11

core (Earth), 353, 353

- density, 335–336, 335, 336, 361t
pressure, 361t
core (Sun)
density, 360t
pressure, 361t
speed distribution of photons in, 517
corona discharge, 645
correspondence principle, 1089, 1106
cosine, 43, 43
cosine-squared rule, for intensity of transmitted polarized light, 902
Cosmic Background Explorer (COBE) satellite, 1242
cosmic background radiation, 1239, 1241, 1242–1243, 1243
cosmic ray protons, 577
cosmic rays, 701
Coulomb barrier, 1207
coulomb per second, 683
Coulomb's law, 565–572, 565 and Gauss' law, 612
spherical conductors, 566, 566–570
coulomb (unit), 565
Crab nebula pulsar, 268
critical angle, for total internal reflection, 911
crossed fields
discovery of electron, 740–741
Hall effect, 741–744, 741, 750
crossed sheets, polarizers, 903, 903
cross product, 49–51, 51
crust (Earth), 353, 353
density, 335, 360t
crystal defects, 577
crystalline solids
electrical properties, 1142, 1142–1143
energy levels, 1143, 1143
crystal planes, 1011, 1011
crystals
matter waves incident after x ray and electron scattering, 1068, 1068, 1069
polycrystalline solids, 880
x-ray diffraction, 1011, 1011–1013
curie (unit), 1175
Curie constant, 878
Curie's law, 878
Curie temperature, 879
currency, anti-counterfeiting measures, 958
current, *See* electric current
current amplitude
series *RLC* circuits, 842–844, 843
current-carrying wire
magnetic field due to, 764–769, 765
magnetic field inside long straight, 773, 773–774
magnetic field outside long straight, 772, 772–773
magnetic force, 750, 750–751
magnetic force between parallel, 770, 770–771
current density, 685–688, 685
current law, Kirchoff's, 714
current-length element, 764, 764
current loops, 683, 683
electrons, 873–874, 873, 874
Faraday's law of induction, 792–794
Lenz's law for finding direction of current, 794, 794–797, 795
as magnetic dipoles, 778–780, 779
solenoids and toroids, 774–778, 775
torque on, 752, 752–753
curvature, of space, 347–348, 347, 1241
cutoff frequency, photoelectric effect, 1060
cutoff wavelength
continuous x-ray spectrum, 1127
photoelectric effect, 1060
cycle
engines, 543
simple harmonic motion, 387
thermodynamic, 489, 490, 492
cyclotrons, 748, 748–749
cylinders
rotational inertia, 255t
tracer study of flow around, 371
calculating capacitance, 660, 660–661
cylindrical symmetry, Gauss' law, 615, 615–616
- D**
damped oscillations
RLC circuits, 833–834
damped simple harmonic motion, 400, 400–401, 401
damped simple harmonic oscillator, 400, 400
damping constant, simple harmonic motion, 400
damping force, simple harmonic motion, 400
dark bands, *See* dark fringes
dark energy, 1243
dark fringes
double-slit interference, 964, 964–965, 965, 967
single-slit diffraction, 992, 992–994, 998
dark matter, 1240
daughter nuclei, 572, 1189
dc (direct current), 705, 835
de Broglie wavelength, 1067, 1085
decay, *See* radioactive decay
decay constant, 1174
decay rate, 1174–1177
deceleration, 20
decibel, 454–456
dees, 748
de-excitation, of electrons to lower energy levels, 1086
deformation, 315, 315
degenerate energy levels, 1095
degrees of freedom, ideal gas molecules, 523–524
density
defined, 7
fluids, 359–361
linear density of stretched string, 420
selected engineering materials, 317t
selected materials and objects, 360t
uniform, for solid bodies, 203–204
density gradient, 1155
density of occupied states, 1149, 1149–1150
density of states, 1146–1147, 1147
depletion zone
MOSFET, 1160
p-n junctions, 1155
derived units, 2
deuterium, 1182
heavy water, 1217
deuterium-tritium fuel pellets, 1212, 1212
deuterons, 1190, 1211
in cyclotron, 749
deuteron-triton reaction, 1211
diamagnetic material, 875
diamagnetism, 875–876, 876
diamond
as insulator, 1144, 1151
unit cell, 1142, 1142
diamond lattice, 1142, 1142
diatomic molecules, 520
degrees of freedom, 523, 523–524, 524t
molar specific heats at constant volume, 520t
potential energy, 193
dielectric constant, 669
selected materials, 669t
dielectrics
atomic view, 671, 671
with capacitors, 669–670
and Gauss' law, 672–674, 672
polarization of light by reflection, 912
dielectric strength, 669–670
selected materials, 669t
differential equation, 830
diffraction. *See also* interference
circular aperture, 1000, 1000–1002
double-slit, 1003, 1003–1005, 1004
Fresnel bright spot, 991, 991
intensity in double-slit, 1003, 1003–1004
intensity in single-slit, 997–999, 998
interference compared and contrasted, 1005
by organized layers, 1011–1013, 1012
single-slit, 963–964, 964, 992, 992–995, 993
and wave theory of light, 990–991
x rays, 1011–1013, 1012
diffraction factor, 1004
diffraction gratings, 1006, 1006–1008
dispersion, 1009, 1009–1011
resolving power, 1009–1011, 1010
three-slit, 1019
x rays, 1011
diffraction patterns
defined, 990–991
double-slit, 1003, 1004
single-slit, 1003, 1004
diffusion current, *p-n* junctions, 1155
dimensional analysis, 420
dip meter, 870
dip north pole, 870
dipole antenna, 890–893, 891
dipole axis, 584, 870
dip-slip, 57
direct current (dc), 835
direct-current (dc) circuits, 705, 835

direction
 of acceleration in one-dimensional motion, 19
 of acceleration in two- and three-dimensional motion, 62–63
 of displacement in one-dimensional motion, 13–14
 of vector components, 41
 of vectors, 39–40, 39, 40
 of velocity in one-dimensional motion, 16
 of velocity in two- and three-dimensional motion, 60
 discharging
 capacitors, 658
 capacitors in *RC* circuits, 720, 720–724
 a charged object, 563
 disintegration, 1169
 disintegration constant, 1174
 disintegration energy, 1177
 disks
 diffraction by circular aperture, 1000, 1000–1003
 rotational inertia, 251t
 dispersion
 chromatic, 907–908, 908
 by diffraction gratings, 1009, 1009–1011
 displacement. *See also* angular displacement; work
 damped harmonic oscillator, 400–401, 401
 one-dimensional motion, 14
 simple harmonic motion, 388, 388, 389
 traveling waves, 417
 two- and three-dimensional motion, 58–60, 59
 as vector quantity, 14, 38
 waves on vibrating string, 414–417, 415
 displacement amplitude
 forced oscillations, 402, 402
 sound waves, 449, 449
 displacement current, 866–868, 867
 displacement vector, 38, 38
 dissipated energy, in resistors, 695, 707
 ac current circuits, 847–848
 distortion parameter, 1200
 distribution of molecular speeds, 516–519, 517
 diverging lens, 936–937, 937
 donor atoms, 1152–1153
 doped semiconductors, 1152, 1152–1154

doping, of semiconductors, 696, 1152–1154
 Doppler effect, 461–465, 1025
 detector moving, source stationary, 462, 462
 for light, 1040–1042, 1041
 source moving, detector stationary, 463, 463
 dose equivalent, radiation, 1184
 dot product, 48, 48
 double-slit diffraction, 1003, 1003–1005, 1004
 double-slit interference
 intensity, 969–972, 970
 from matter waves, 1067–1068, 1068
 single-photon, wide-angle version, 1067, 1067
 single-photon version, 1066
 Young's experiment, 964, 964–968
 doubly magic nuclides, 1186
 down quark, 578, 655, 1232–1234, 1234t
 drag coefficient, 122
 drag force, 121–123
 damped simple harmonic motion, 400
 mechanical energy not conserved in presence of, 175
 as nonconservative force, 168
 drain, FETs, 1159, 1159
 drift current, *p-n* junctions, 1155
 drift speed, 686–688
 Hall effect for determining, 741, 741–742
 driven oscillations, 402, 836, 836
 driving angular frequency, 835
 driving frequency, of emf, 835
d subshells, 1125, 1127

E
 E (exponent of 10), 2
 Earth, 330, 1243. *See also* core; crust; gravitational force; mantle
 angular momentum effects
 if polar ice caps melted, 304
 atmospheric electric field, 580t, 656
 average density, 360t
 density as function of distance from center, 335
 eccentricity of orbit, 343

effective magnetic dipole moment, 1117
 ellipsoidal shape of, 335
 escape speed, 341t, 341–342
 gravitation near surface, 334–337, 336
 interior of, 353, 353
 Kepler's law of periods, 344t
 level of compensation, 380
 magnetic dipole moment, 754, 754t, 887
 magnetism, 869–871
 nonuniform distribution of mass, 335, 335–337
 rotation, 336, 336
 satellite orbits and energy, 345–346, 345
 variation in length of day over 4-year period, 6

earthquakes
 building oscillations during, 386
 buildings submerged during, 7
 natural angular frequency of buildings, 402, 402
 S and P waves, 468

Earth's magnetic field, 738, 870, 870–871
 magnetic latitude, 888
 polarity reversal, 870, 870–871
 at surface, 738t

eccentricity, of orbits, 342–343, 343
 and orbital energy, 345
 planets of Solar System, 344t

eddy currents, 800
 edge effect, 618
 effective cross-sectional area, 122
 effective magnetic dipole moment, 1117
 effective phase difference, optical interference, 962

efficiency
 Carnot engines, 545–546
 real engines, 546, 549–550
 Stirling engines, 546

eightfold way, 1231, 1231t, 1231–1232

Einstein, Albert, 87, 893, 1022, 1022, 1025, 1043. *See also* relativity

Bose-Einstein condensate, 1221, 1221

and bosons, 1221
 and lasers, 1132
 view of gravitation, 347, 347–348

work on photoelectric effect, 1059–1062
 work on photons, 1057–1059

Einstein-de Haas experiment, 1114, 1114

Einstein ring, 348, 348

elastic bodies, 315

elastic collisions
 defined, 217
 in one dimension with moving target, 222–223, 223

in one dimension with stationary target, 221, 221–222

in two dimensions, 224, 224

elasticity, 305, 315–318, 316

and wave speed on stretched string, 420

elastic potential energy, 167
 determining, 171
 traveling wave on stretched string, 421–422, 421

electrical breakdown, 580t, 593, 593

electrically isolated object, 562

electrically neutral objects, 562

electrical-mechanical analogy, 829t, 829

electric charge, 561–562. *See also* circuits conductors and insulators, 563–564
 conservation of, 572–573
 and Coulomb's law, 565–572, 565

induced, 564
 LC oscillations, 826, 830
 measures of, 586, 586t
 quantization, 571

electric circuits, *See* circuits

electric current, 682–685, 683
 calculating in single-loop circuits, 707, 707–709
 current density, 685–688, 685
 direction in circuits, 684, 684–685

induced, 791, 797–800
 LC oscillations, 826, 831

magnetic field due to, 764–769, 765

time-varying in *RC* circuits, 720–724

electric dipole, 582, 754
 in electric field, 594–596
 electric field due to, 584, 584–585

electric potential due to, 637–638, 638

induced, 638

potential energy, 594–595

electric dipole antenna, 891, 891–892

- electric dipole moment, 585, 585–594
 dielectrics, 671
 permanent, 638
 electric displacement, 673
 electric field, 580–581, 736
 calculating from potential, 641, 641–642
 calculating potential from, 633, 633–634
 capacitors, 659
 crossed fields, 741–744, 741
 and current, 682
 due to charged disk, 591, 640, 640
 due to electric dipole, 584, 584–585
 due to line of charge, 586–590, 586
 due to point charge, 582, 582–583
 electric dipole in, 594–596
 energy stored in capacitor, 667–669
 equipotential surfaces, 631, 631–632, 632
 flux, 606, 606–608, 884
 Gauss' law for, 605, 608–620, 771, 862, 869t
 Hall effect, 741–744, 741, 750
 induced, 800–804, 801
 LC oscillations, 826
 point charge in, 592–593
 polarized light, 902, 902–904, 903
 system of point charges, 642–644, 643
 traveling electromagnetic waves, 890–893, 891, 893
 work done by, 628
 electric field lines, 581, 581–582
 electric fish, 718
 electric flux, 606, 606–608, 884
 electric force, 562, 735
 electric generator, 705
 electric motors, 752, 752–753, 869
 electric potential, 629–631
 calculating field from, 641, 641–642
 calculating from field, 633, 633–634
 charged isolated sphere, 645, 645
 due to charged disk, 640, 640
 due to continuous charge distribution, 639, 639–640, 640
 due to electric dipole, 637–638, 638
 due to group of point charges, 636, 636–637
 due to line of charge, 639, 639–640
 due to point charge, 635, 635
 and induced electric field, 803–804
 electric potential difference, 630
 electric potential energy, 628–629, 629
 system of point charges, 642–644, 643
 electric quadrupole, 600, 600
 electric spark, 593, 593
 airborne dust explosions set off by, 668–669
 dangers of, 645, 645
 and pit stop fuel dispenser fire, 723, 723–724
 electric wave component, of electromagnetic waves, 891–893, 893
 electromagnetic energy, 832
 electromagnetic force, 1221, 1235–1236
 messenger particle, 1235–1237
 electromagnetic oscillations, 826–829, 827
 damped in RLC circuits, 833–834
 forced, 836, 836–838
 electromagnetic radiation, 890
 electromagnetic spectrum, 889–890, 890
 electromagnetic waves, 413, 889–890. *See also reflection; refraction*
 energy transport, 897–899, 898
 induced electric field, 894, 894–896
 induced magnetic field, 895, 895–896
 intensity variation with distance, 898, 898–899
 polarization, 901, 901–904, 902, 903
 polarization by reflection, 912, 912–913
 Poynting vector, 897–899, 898
 radiation pressure, 899–900
 speed of travel, 413, 893
 traveling, 890–896, 893, 894
 electromagnetism, 764, 869
 electromagnets, 735, 735
 magnetic field near big, 738t
 electromotive force (emf), 705.
See also emf
 electron capture, 1180
 electron diffraction, 1069
 electron microscope, 1080
 electron neutrinos, 1227, 1227t
 electron-positron annihilation, 572, 572, 1222
 electrons, 563, 1219
 accelerator studies, 748
 barrier tunneling, 1074, 1074–1076, 1075
 Bohr model, 1096–1099, 1097
 bubble chamber tracks, 572, 572, 738
 charge, 571, 571t
 Compton scattering, 1063, 1063–1065, 1063
 conduction, 1145–1150
 discovery by Thomson, 740, 740–741, 1165
 as fermions, 1220
 kinetic energy dependence on speed, 1023, 1023
 as leptons, 1221, 1227, 1227t
 loop model for orbits, 873–874, 873, 874
 magnetic dipole moment, 754, 754t
 and magnetism, 869–874
 majority carrier in *n*-type semiconductors, 1153, 1153t
 matter waves, 1067–1070, 1068, 1069, 1083–1084
 minority carrier in *p*-type semiconductors, 1153, 1153t
 orbital angular momentum, 872–873, 873
 from proton-antiproton annihilation, 1224t
 radiation dosage, 1184
 spin, 1115, 1115t, 1220–1221, 1221
 spin angular momentum, 871–872, 872
 spin-flip, 883
 in superconductors, 697–698
 electron spin, 1115, 1115t, 1220–1221, 1221
 electron traps, 1115
 energies of, 1084–1088
 finite well, 1091, 1091–1093
 multiple electrons in rectangular, 1121–1124
 nanocrystallites, 1093, 1093
 one-dimensional, 1084–1095
 quantum corrals, 1094, 1094
 quantum dots, 1083, 1094, 1094
 two- and three-dimensional, 1095, 1095–1096
 wave functions, 1088–1091, 1089
 electron-volt, 630
 electroplaques, 718, 718–719
 electrostatic constant, 565
 electrostatic equilibrium, 613
 electrostatic force, 562, 580
 and Coulomb's law, 565–572, 565
 electric field due to point charge, 582, 582–583
 point charge in electric field, 592–593
 work done by, 628, 630–631
 electroweak force, 1236
 elementary charge, 570
 measuring, 592
 emf
 alternating current, 835
 defined, 705
 and energy and work, 706, 706–707
 induced, 792
 potential and power in circuits, 711–714
 self-induced, 806
 emf devices, 705. *See also batteries*
 internal dissipation rate, 712
 real and ideal, 706–706, 707
 emf rule, 709
 emission event, photons, 1058
 lasers, 1132, 1132–1133
 emission lines, 1006, 1006, 1100–1101
 emissivity, 496
 endothermic reactions, 1227
 energy. *See also gravitational potential energy; kinetic energy; mechanical energy; potential energy; thermal energy; work*
 conservation of, 140, 183–186, 184
 dark energy, 1243
 defined, 140
 and relativity, 1043–1047
 scalar nature of, 38
 in simple harmonic motion, 392–394, 393
 as state property, 537
 stored in electric field, 667–669
 stored in magnetic field, 811–812
 transport by electromagnetic waves, 897–899, 898
 traveling wave on stretched string, 421, 421–423
 energy bands, 1143–1144, 1144

- energy density
capacitors, 668
electric field, 668
kinetic energy density of fluids, 374
magnetic field, 812–813
energy gap, 1143–1144, 1144
energy-level diagrams, 1085, 1085
one-electron, 1123, 1123
energy levels
excitation and de-excitation, 1086
full, empty, and partially occupied, 1122
hydrogen, 1099–1100, 1099, 1100
multiple electron traps, 1121–1124
nuclear, 1172
single electron traps, 1084–1088, 1085
energy method, of calculating current in single-loop circuits, 707
engines
Carnot, 543–546, 543, 549–550
efficiency, 543, 545–546, 549, 549–550, 550
ideal, 543–544
perfect, 546, 546
Stirling, 546, 546
entoptic halos, 1014, 1016
entropy
engines, 543–546
force due to, 542
and irreversible processes, 536
and probability, 552–553
refrigerators, 548–549
and second law of thermodynamics, 541–542
as state function, 537, 539–540
statistical mechanics view of, 550–553
entropy changes, 536, 537–541
Carnot engines, 545
Stirling engines, 545
entropy postulate, 536
equation of continuity, 372–374, 372
equations of motion
constant acceleration, 23, 23t
constant linear and angular acceleration contrasted, 248t
free-fall, 25
equilibrium, 92. *See also* static equilibrium
defined, 305–306
electrostatic, 613
examples of static, 309–313
indeterminate structures, 314, 314
protons, 569
requirements of, 306–307
secular, 1191
thermal, 477
equilibrium charge, capacitors in *RC* circuits, 720
equilibrium points, in potential energy curves, 176–177
equilibrium position, simple pendulum, 395
equilibrium separation, atoms in diatomic molecules, 193
equipartition of energy, 523
equipotential surfaces, 631, 631–633
equivalence, of gravity and acceleration, 347
equivalent capacitance/capacitor
parallel capacitances, 662–663, 663, 715t
series capacitances, 663–666, 715t
equivalent inductance/inductor
parallel inductances, 822
series inductances, 822
equivalent resistance/resistor
parallel resistances, 714, 714–719, 715t
series resistances, 710, 715t
escape speed, 341–342
selected astronomical bodies, 341t
event horizon, 337
events, 1022
Lorentz factor, 1028, 1028, 1029, 1043
Lorentz transformation, 1035, 1035–1036
measuring, 1024–1025, 1025
relativistic reversal, 1051
relativity of length, 1031, 1031–1034, 1038
relativity of simultaneity, 1025–1026, 1026, 1037
relativity of time, 1027, 1027–1031, 1037
relativity of velocities, 1039, 1039
exchange coupling, 879, 887
excitation, of electrons, 1086, 1086
excitation energy, 1110
excited states, 1085, 1086
expansion, of universe, 1238, 1242–1243, 1243
explosions
one-dimensional, 216, 216–217
two-dimensional, 216–217, 217
extended objects, 100, 927, 927
locating images by drawing rays, 939, 939
external electric field, 591
Gaussian surfaces, 612, 612–614
isolated conductor in, 645, 645
external forces, 92
collisions and internal energy transfers, 184–185
system of particles, 206–209
work done with friction, 180–182
work done without friction, 180
external magnetic field
and diamagnetism, 875–876
and ferromagnetism, 875, 879–880, 879
and paramagnetism, 875, 877–878
external torque, 288
and conservation of angular momentum, 290, 291
eye, *See* human eye
eyepiece
compound microscope, 944, 944
refracting telescope, 945
- F**
face-centered cubic, 1142, 1142
Fahrenheit temperature scale, 479–481, 480
farad, 657
Faraday, Michael, 561, 581, 669, 791, 792, 805
Faraday's law of induction, 792–794, 863
Faraday's experiments, 792
and Lenz's law, 794, 794–797, 795
Maxwell's equation form, 869t
mutual induction, 814
reformulation, 802
self-induction, 806, 806–807
transformers, 850
femtometer, 1170
fermi (unit), 1170
Fermi, Enrico, 1196, 1206, 1220
Fermi-Dirac statistics, 1148
Fermi energy, 1145, 1146, 1148
calculating, 1150
effect of doping on, 1163–1164, 1164
Fermilab accelerator, 1219, 1234
Fermi level, 1145
fermions, 1220–1221, 1221
quarks as, 1233t
Fermi speed, 1145
ferromagnetic materials, 875, 879–880, 879, 880
magnetically hard and soft, 883
ferromagnetism, 875, 879–880, 879. *See also* iron
FET (field-effect-transistor), 1159, 1159
field declination, 870
field-effect-transistor (FET), 1159, 1159
field inclination, 870
field of view
refracting telescope, 945
spherical mirrors, 928
final state, 488, 489, 520
finite well electron traps, 1091, 1091–1092
fires
pit stop fuel dispenser fire, 723, 723–724
first law of thermodynamics, 491
special cases of, 492t, 492–493
first-order line, 1006
first reflection point, 921, 921
fish, electric, 718–719
fission fragments, 1196–1198
fixed axis, 241
angular momentum of rigid body rotation about, 288–290, 289
flavors, of quarks, 1232, 1236–1237
floaters, 991, 1017
floating, 367, 368
fluids, 121
apparent weight in, 369–370
Archimedes' principle, 367, 367–370
Bernoulli's equation, 374–377
defined, 359
density, 359–360
equation of continuity, 372, 372–373
hydraulic lever, 366, 366–367
motion of ideal, 371–372, 372
Pascal's principle, 366, 366–367
pressure, 360–361

- pressure measurement, 365–367, 366
at rest, 362, 362–364
fluid streamlines, 371–372, 372
flux, 605–606, 606
of electric field, 606,
606–608, 884
of magnetic field, 862
focal length
compound microscope, 944,
944–945
refracting telescope, 945, 945
simple magnifying lens, 943,
943–944
spherical mirrors, 928,
928–930
thin lenses, 936–937, 937
focal plane, 967
focal point
compound microscope, 944,
944–945
refracting telescope, 945,
945–946
simple magnifying lens, 943,
943–944
spherical mirrors, 928,
928–929
thin lenses, 936–937, 937
two-lens system, 939–940,
940
- focus, *See* focal point
force constant, 149
forced oscillations, 402, 836,
836
forced oscillators, 402, 402
forces, 289t. *See also* acceleration; angular momentum; collisions; drag force; electrostatic force; external forces; friction; gravitational force; internal forces; linear momentum; magnetic force; spring force; torque; variable force; work
attractive, 331
buoyant, 367, 367–370
centripetal, 124–128, 125
conservative, 168–170, 169,
629
defined, 87
due to entropy, 542
equilibrium, 92
and motion, 13
net, 92
and Newton's first law, 87–91
Newton's laws applications,
100–105
and Newton's second law,
91–94
- and Newton's third law, 98
nonconservative, 168
normal, 96–97, 97
path independence of conservative, 168–170, 169
principle of superposition for, 88
resultant, 92
of rolling, 278–280, 279
tension, 97–98, 98
unit of, 88, 88–90
as vector quantities, 88
and weight, 95–96
- forward-bias connection, junction rectifiers, 1156–1157, 1156, 1157
- Franklin, Benjamin, 562, 570,
572, 701
- Fraunhofer lines, 1141
- free-body diagrams, 92–94, 93
Newton's laws applications, 100–105
free charge, 673
free-electron model, 1115,
1145
of Ohm's law, 693
free electrons, 682
free expansion
first law of thermodynamics for, 492t, 492
ideal gases, 527–528, 537,
537–541, 538
- free-fall acceleration (g),
25–26, 25
measuring with physical pendulum, 397
- free-fall flight, 25
- free oscillations, 402, 836
- free particle, 1071
and Heisenberg's uncertainty principle, 1072,
1073
matter waves, 1084
- free space, 890
- freezing point, 486
- frequency, 414–417
circulating charged particle,
744–745
and index of refraction, 961
photons, 1058
simple harmonic motion,
386–388, 388
- sound waves, 449
wave on stretched string, 421
- waves, 418, 418
- Fresnel bright spot, 991, 991
- friction, 97, 97, 117, 116–121.
See also kinetic frictional force; static frictional force
cold-weld, 118, 118, 119
- as nonconservative force (kinetic friction), 168
and rolling, 278–279, 279
work done by external force with, 180–182, 180, 181
- frictionless surface, 88, 97
- fringing, 618
- fuel charge, nuclear reactor, 1205
- fuel rods, 1202, 1205
- fulcrum, 321
- full levels, 1122
- full width at half maximum, 1016
- fully charged capacitor, 658
- fully constructive interference, 426, 427, 427t, 431
sound waves, 451
- fully destructive interference, 426, 427, 427t, 431
sound waves, 451
- fundamental mode, 434, 457
- fur
rubbing plastic rod with, 562
- fused quartz
coefficient of linear expansion, 482t
- index of refraction, 906t
- index of refraction as function of wavelength, 907
- resistivity, 689t
- fusion, 1045, 1172, 1207–1208
controlled, 1211–1213
in Sun and stars, 1207,
1209–1210, 1209
- fusion reaction, 1045
- G**
- g* units, of acceleration, 20
g, free-fall acceleration, 25–26
measuring with physical pendulum, 397
- G*, gravitational constant, 331
- galaxies, 330
Doppler shift, 1041
Doppler shift of NGC 7319,
1052, 1052
- formation in early universe,
1241
gravitational lensing caused by, 348, 348
- preponderance of matter over antimatter, 1222
- recession, and expansion of universe, 1238
- superluminal jets, 1054
- Galilean transformation equations, 1035
- Galileo, 355
- gamma-ray photons, 1222
in fusion, 1209
- gamma rays, 573, 737, 890
bubble chamber track, 1069,
1069
- radiation dosage, 1184
- ultimate speed, 1023
- gas constant, 508
- gases, 507. *See also* ideal gases; kinetic theory of gases
compressibility, 360
confined to cylinder with movable piston,
488–490, 489
- density of selected, 360t
as fluids, 359
specific heats of selected, 485t
speed of sound in, 447t
thermal conductivity of selected, 494t
- gas state, 486
- gauge (wire), 700
- gauge pressure, 363
- gauss (unit), 738
- Gauss, Carl Friedrich, 605
- Gaussian form, of thin-lens formula, 956
- Gaussian surfaces
capacitors, 659–662
defined, 605, 605
electric field flux through,
606, 606–608
external electric field,
613–614, 614
and Gauss' law for magnetic fields, 862
- Gauss' law, 605
charged isolated conductor, 612–615
conducting plates, 617,
617–618
and Coulomb's law, 612
cylindrical symmetry, 615,
615–616
dielectrics, 672–674, 672
for electric fields, 605,
608–620, 862, 869t
isolated conductor with cavity, 612, 613
for magnetic fields, 862–863,
863, 869t
Maxwell's equation forms,
869t
nonconducting sheet, 617,
617
planar symmetry, 617–618,
617
spherical symmetry,
619–620, 619
- Geiger counter, 627, 627, 1165
- general theory of relativity,
347–348, 1022, 1029

geomagnetic pole, 738, 870
 geometrical optics, 904, 990
 diffraction as limit to, 963–964
 geosynchronous orbit, 354
 glass. *See also* Pyrex glass
 coefficient of linear expansion, 482t
 index of refraction, 906t
 as insulator, 563
 polarization of light by reflection, 912
 rubbing rod with silk, 562, 562, 563
 shattering by sound waves, 454
 Global Positioning System (GPS), 1, 1022
 g-LOC (*g*-induced loss of consciousness), 72
 in dogfights, 380
 gluons, 747, 1233, 1236
 gold, 1129
 alpha particle scattering, 1166
 impact with alpha particle, 644
 isotopes, 1169
 grand jeté, 208, 208
 grand unification theories (GUTs), 1237
 graphical integration, for one-dimensional motion, 27, 27–28
 grating spacing, 1006–1007
 grating spectroscope, 1008, 1008
 gravitation, 330
 and Big Bang, 1241
 Einstein's view of, 347–348, 348
 inside Earth, 337–338
 near Earth's surface, 334–337, 335
 Newton's law of, 330–332, 331, 342
 gravitational acceleration (a_g), 334
 variation with altitude, 334t
 gravitational constant (G), 331
 gravitational force, 94, 1221
 center of gravity, 308, 308–309
 electrostatic force contrasted, 572
 and Newton's law of gravitation, 330–332, 331
 pendulums, 395, 395
 and potential energy, 340–341

and principle of superposition, 333–334
 work done by, 146, 146–148
 gravitational lensing, 348, 348
 gravitational potential energy, 166, 339–342, 339
 determining, 171
 and escape speed, 341–342
 and gravitational force, 340–341
 graviton, 348
 gray (unit), 1184
 grounding, 563
 ground speed, 84
 ground state, 1085, 1086, 1086
 wave function of hydrogen, 1102–1103, 1103
 zero-point energy, 1090
 g subshells, 1125
 gyroscope precession, 294–295, 295

H

hadrons, 1221, 1228–1230
 half-life, 1170, 1175, 1219
 and radioactive dating, 1183
 half-width of diffraction grating lines, 1006–1007, 1007
 resonance curves, 845–846, 846
 Hall effect, 741–744, 741, 750, 1142
 Hall potential difference, 742
 halo nuclides, 1170
 hanging blocks, 100, 100–101
 hard reflection, of traveling waves at boundary, 432
 harmonic motion, 387
 harmonic number, 434
 musical sounds, 456–459
 harmonic series, 434
 hearing threshold, 454t
 heat, 547–548. *See also* thermodynamics
 absorption by solids and liquids, 485–488
 defined, 484
 first law of thermodynamics, 491
 path-dependent quantity, 491
 signs for, 483–484
 and temperature, 483–484
 heat capacity, 485
 heat engines, 543–548
 heat of fusion, 487
 selected substances, 487t
 heat of vaporization, 486
 selected substances, 487t
 heat pumps, 548
 heats of transformation, 486–488
 selected substances, 487t
 heat transfer, 494–497
 Heisenberg's uncertainty principle, 1072, 1073, 1235–1236
 helical paths, charged particles, 745–746, 746
 helium-neon gas laser, 1133, 1133–1135
 Helmholtz coil, 787, 788, 790
 helium-burning, in fusion, 1210
 henry (unit), 805
 hertz, 386
 holes, 1128, 1151
 majority carrier in *p*-type semiconductors, 1153, 1153t
 minority carrier in *n*-type semiconductors, 1153, 1153t
 holograms, 1131
 home-base level, for spectral series, 1101
 Hooke, Robert, 149
 Hooke's law, 149, 176
 hoop
 rotational inertia, 255t
 horizontal motion, in projectile motion, 66–67, 67
 horizontal range, in projectile motion, 65, 66
 horsepower (hp), 155
 h subshells, 1125
 Hubble constant, 1238
 Hubble's law, 1238
 human body
 as conductor, 563
 physiological emf devices, 706
 human eye, 943, 955, 955
 floaters, 991, 1017
 image production, 926
 sensitivity to different wavelengths, 890, 890
 Huygens' principle, 959, 959–960
 Huygens' wavelets, 959, 992, 995
 hydraulic compression, 317
 hydraulic engineering, 359
 hydraulic jack, 367
 hydraulic lever, 366, 366–367
 hydraulic stress, 316, 316–318
 hydrogen
 Bohr model, 1096–1099, 1097
 electric field within, 580t
 emission lines, 1008, 1008

I

ideal emf devices, 706
 ideal engines, 543–544
 ideal fluids
 motion, 371–372, 372
 ideal gases, 508–511
 adiabatic expansion, 526–528
 average speed of molecules, 516–519
 free expansion, 537, 537–541, 538
 internal energy, 519–522
 molar specific heat and degrees of freedom, 523–524
 molar specific heat at constant pressure, 521–522
 molar specific heat at constant volume, 519–520
 molar specific heats, 519–522
 most probable speed of molecules, 517
 RMS speed, 511, 511–513, 512t, 516
 translational kinetic energy, 513–514
 work done at constant pressure, 510
 work done at constant temperature, 509, 509–510
 work done at constant volume, 510
 ideal gas law, 508
 ideal gas temperature, 479
 ideal inductor, 807
 ideal refrigerators, 548

- ideal solenoid, 776–777
 ideal spring, 149
 ideal toroids, 778
 ideal transformers, 850,
 850–852
 ignition, in magnetic confinement, 1212
 image distances, 926
 images
 extended objects, 927, 927,
 939, 939
 locating by drawing rays,
 932, 932–933, 939, 939
 from plane mirrors, 926,
 926–928
 from spherical mirrors, 928,
 928, 930, 930–934, 946,
 946
 from spherical refracting surfaces, 934, 934–935,
 946–947, 946
 from thin lenses, 938,
 938–942, 939, 947–948,
 947
 types of, 924–925
 impedance, 843
 matching in transformers,
 852
 impulse, 212
 series of collisions, 213, 213
 single collision, 211, 211–212
 incident ray, 904–906, 905
 incoherent light, 969
 incompressible flow, 371
 indefinite integral, 24
 independent particle model, of nucleus, 1185–1186
 indeterminate structures
 equilibrium, 314, 314
 index of refraction
 and chromatic dispersion, 907–908
 common materials, 906t
 defined, 906, 960
 and wavelength, 960–962
 induced charge, 564
 induced current, 791
 induced electric dipole moment, 638
 induced electric fields, 800–804, 801
 from traveling electromagnetic waves, 894,
 894–896
 induced emf, 792
 induced magnetic fields, 863–866, 864
 displacement current, 866–869, 867
 finding, 867–868, 868
 from traveling electromagnetic waves, 895,
 895–896
 inductance, 805–806
 LC oscillations, 826–829
 RL circuits, 807–810
 RLC circuits, 833–834
 solenoids, 805–806, 806
 induction, 792. *See also*
 Faraday's law of induction
 energy transfer, 797–800, 798
 mutual induction, 813–815,
 814
 self-induction, 806, 806–807,
 813
 inductive reactance, 841
 inductive time constant, 808–809
 inductors, 805, 805–807
 with ac generator, 840,
 840–841, 841
 LC oscillations, 826–829, 827
 in parallel, 822
 phase and amplitude relationships for ac circuits, 842t
 RL circuits, 807–810
 RLC circuits, 833, 833–834
 in series, 822, 844–853
 series *RLC* circuits, 842–844, 843
 inelastic collisions
 defined, 218
 in one dimension, 218,
 218–220, 219
 in two dimensions, 224
 inertial confinement, 1212
 inertial reference frames, 89–90, 1022
 inexact differentials, 491
 infinitely deep potential energy well, 1085, 1085
 infinite potential well, 1085
 inflation, of early universe, 1241
 infrared radiation, 899
 initial conditions, 391
 initial state, 488, 489, 519
 in phase
 ac circuits, 842t
 resistive load, 838
 sound waves, 451
 thin-film interference, 973–974
 traveling electromagnetic waves, 890–891
 waves, 425, 428
 instantaneous acceleration
 one-dimensional motion, 18
 two- and three-dimensional motion, 62–63
 instantaneous angular acceleration, 243
 instantaneous angular velocity, 243
 instantaneous power, 155, 185
 instantaneous velocity
 one-dimensional motion, 17–18
 two- and three-dimensional motion, 60
 insulators, 563–564, 697
 electrical properties, 1144,
 1144
 resistivities of selected, 689t
 unit cell, 1142
 integrated circuits, 1159–1160
 intensity
 diffraction gratings, 1006
 double-slit diffraction, 1003,
 1003–1005
 double-slit interference, 969–972, 970
 electromagnetic waves, 898,
 898–899
 single-slit diffraction, 995–999, 996, 998
 of transmitted polarized light, 902–904, 903
 intensity, of sound waves, 452–456, 453
 interference, 425t, 425–428,
 426. *See also* diffraction and coherence, 969
 combining more than two waves, 972
 diffraction compared and contrasted, 1003–1004
 double-slit from matter waves, 1067–1069,
 1068
 double-slit from single photons, 1065, 1065–1067
 intensity in double-slit, 969–972, 970
 Michelson's interferometer, 980, 980
 Newton's rings, 987, 987
 and rainbows, 962, 962
 sound waves, 451, 451–452
 thin films, 973, 973–979
 Young's double-slit experiment, 964, 964–968, 965
 interference factor, 1004
 interference fringes, 965, 965
 locating, 965, 965–967
 interference pattern, 964, 964,
 967
 interfering waves, 425
 interferometer, 980
 intermediate interference, 426,
 427, 427t
 sound waves, 451
 internal energy, 476
 and conservation of total energy, 183
 and external forces, 184–185
 and first law of thermodynamics, 491
 of ideal gas by kinetic theory, 519–522
 internal forces, 92
 system of particles, 206–9
 internal resistance
 ammeters, 720
 circuits, 709, 709
 emf devices, 712–713
 internal torque, 288
 International Bureau of Weights and Standards, 3, 6
 International System of Units, 2
 interplanar spacing, 1012
 intrinsic angular momentum electrons, 871
 protons and neutrons, 872
 inverse cosine, 43, 43
 inverse sine, 43, 43
 inverse tangent, 43, 43
 inverse trig functions, 43, 43
 inverted images, 930, 930
 ionization energy, 1113, 1113
 plot of elements, 1113
 ionized atoms, 1101
 iron, 1126–1127. *See also* steel
 Curie temperature, 879
 ferromagnetic material, 875,
 879–880
 quantum corral, 1094, 1094
 radius of nucleus, 571
 resistivity, 689t
 iron filings
 bar magnet's effect on, 862,
 862
 current-carrying wire's effect on, 765
 irreversible processes, 536. *See also* entropy and second law of thermodynamics, 541–542
 irrotational flow, 371, 374
 island of stability, 1170
 isobaric processes summarized, 528, 528t
 isobars, 1170
 isochoric processes summarized, 528, 528t
 isolated spherical capacitors, 669
 calculating capacitance, 661–662

isolated system, 173
conservation of total energy, 183–184
linear momentum conservation, 215
isotherm, 509, 509
isothermal compression, 509
Carnot engine, 543, 543
isothermal expansion, 509
Carnot engine, 543, 543
entropy change, 537–538, 538
isothermal processes
summarized, 528, 528t
isotopes, 1169
isotopic abundance, 1168t
isotropic materials, 690
isotropic point source, 898
isotropic sound source, 454

J

joint, in rock layers, 131
Josephson junction, 1075
joule (J), 141, 485
junction diodes, 697
junction lasers, 1158, 1158
junction plane, 1154, 1155
junction rectifiers, 1156, 1156–1157, 1164
junction rule, Kirchoff's, 714
junctions, circuits, 714

K

kaons, 1030, 1219
and eightfold way, 1231t
and strangeness, 1231
kelvins, 476, 478, 479, 482
Kelvin temperature scale, 476, 476, 480
Kepler, Johannes, 342
Kepler's first law (law of orbits), 342–343, 343
Kepler's second law (law of areas), 343, 343
Kepler's third law (law of periods), 344, 344
for planets in Solar System, 344t
kilogram, 6
kilowatt-hour, 155
kinematics, 13
kinetic energy, 263t. *See also* collisions; rotational kinetic energy; translational kinetic energy; work-kinetic energy theorem
collisions, 217–218
and conservation of mechanical energy, 173–175

and conservation of total energy, 183–186
defined, 141
and momentum, 1046
and relativity, 1045, 1045–1047
of rolling, 277, 277–280
of rotation, 253–254, 254
satellites in orbit, 345–346, 346
simple harmonic motion, 393, 393
traveling wave on stretched string, 421, 421, 422
and work, 142–146, 143
yo-yo, 281
kinetic energy density, of fluids, 374
kinetic energy function, 176
kinetic frictional force, 118, 118
as nonconservative force, 168
rolling wheel, 278
kinetic theory of gases. *See also* gases; ideal gases
average speed of molecules, 516–518
and Avogadro's number, 507–508
distribution of molecular speeds, 516–519, 517
mean free path, 514, 514–515
molar specific heat and degrees of freedom, 523–524
most probable speed of molecules, 517
and quantum theory, 523, 525
RMS speed, 511, 511–513, 512t, 516
translational kinetic energy, 513–514
Kirchoff's current law, 714
Kirchoff's junction rule, 714
Kirchoff's loop rule, 708
Kirchoff's voltage law, 708
K shell, 1128, 1129, 1136

L

lagging, in ac circuits, 841, 842t, 843
lagging waves, 428
lambda particle and eightfold way, 1231t
lambda-zero particle, 1230
laminar flow, 371
lasers, 1131–1132
coherence, 969
helium-neon gas, 1133, 1133–1135

junction lasers, 1158, 1158
nuclear-pumped x-ray, 1017
operation, 1132, 1132–1135
for plasma compression, 916
radiation pressure, 899–900
surgery applications, 1131, 1131
laser thickness monitoring, 1015, 1015
lasing, 1134, 1140
lateral magnification
compound microscope, 944
spherical mirrors, 931
two-lens system, 939
lateral manipulation, using STM, 1075
lattice, 315, 315, 1142, 1142
law of areas (Kepler's second law), 343, 343
law of Biot and Savart, 765–766, 771, 779
law of conservation of angular momentum, 290, 290–294, 291, 292
law of conservation of electric charge, 572–573
law of conservation of energy, 183–186, 184
law of conservation of linear momentum, 215
law of orbits (Kepler's first law), 342–343, 343
law of periods (Kepler's third law), 344, 344
for planets in Solar System, 344t
law of reflection, 906
law of refraction, 906, 959, 959–962
Lawson's criteria, 1211
LC oscillations
and electrical-mechanical analogy, 830t, 830
qualitative aspects, 826–829, 827
quantitative aspects, 830–833
LC oscillators, 830t, 830–833
radio wave creation, 890–893, 891
lead
coefficient of linear expansion, 482t
heats of transformation, 487t
specific heats, 485t
thermal conductivity, 494t
leading, in ac circuits, 841, 842t, 843
leading waves, 428
Leidenfrost effect, 504
length

relativity of, 1031, 1031–1034, 1037
length contraction, 1032–1034
consequences of Lorentz transformation equations, 1037–1038
length units, 3–4
lens, 936. *See also* thin lenses
diffraction by, 1000
lens maker's equation, 936
Lenz's law, 794, 794–797, 795
self-induction, 806
lepton number, 1228
conservation of, 1228
leptons, 1221, 1227t, 1227–1228
conservation of lepton number, 1228
formation in early universe, 1241
lifetime, subatomic particles, 1029
lifting capacity, balloons, 534
light, 413, 894. *See also* diffraction; interference; photons; reflection; refraction
Doppler effect, 461
in early universe, 1241
Huygens' principle, 959, 959–960
law of reflection, 906
law of refraction, 906, 959, 959–962
as probability wave, 1065–1067
travel through media of different indices of refraction, 961, 961
as wave, 959, 959–963
wave theory of, and diffraction, 990–991
light-emitting diodes (LEDs), 1157, 1157–1158, 1158
light-gathering power refracting telescope, 945
lightning, 561, 656
role in creation of lode-stones, 881
side flash, 728, 728
upward streamers, 616, 616
light quantum, 1057–1058
light wave, 893
light-year, 12
line (spectral), 1101
linear acceleration, *See* acceleration
linear charge density, 586, 586t
linear density, of stretched string, 420, 421
linear expansion, 482, 482

- linear momentum, 210, 289t
 completely inelastic collisions in one dimension, 218–219
 conservation of, 215–217
 elastic collisions in one dimension with moving target, 222–223
 elastic collisions in one dimension with stationary target, 221–222
 elastic collisions in two dimensions, 224
 at equilibrium, 305
 and impulse of series of collisions, 213
 and impulse of single collision, 211–212
 inelastic collisions in one dimension, 218, 218–220, 219
 inelastic collisions in two dimensions, 224
 photons, 1062, 1062–1065, 1063
 system of particles, 211
- linear momentum-impulse theorem**, 212
- linear motion**, 241
- linear oscillator**, 390–392
- linear simple harmonic oscillators**, 390, 390–392
- line integral**, 633
- line of action**, of torque, 259, 259
- line of charge**
 electric field due to, 586–590, 586
 electric potential due to, 639, 639–640
- line of symmetry**, center of mass of solid bodies with, 203
- lines**, diffraction gratings, 1006–1007
- lines of force**, 581
- lines shapes**, diffraction gratings, 1010
- liquefaction**, of ground during earthquakes, 7
- liquids**
 compressibility, 317, 360
 density of selected, 360t
 as fluids, 359
 heat absorption, 485–488
 speed of sound in, 447t
 thermal expansion, 481–483
- liquid state**, 486
- Local Group**, 330
- Local Supercluster**, 330
- lodestones**, 561, 869, 881
- longitudinal motion, 414
- longitudinal waves, 413–414, 414
- long jump**
 conservation of angular momentum, 291, 291
- long wave radiation**, 890
- loop model**, for electron orbits, 873–874, 873, 874
- loop rule**, 708, 714
RL circuits, 808, 808
- Lorentz factor**, 1028, 1028, 1029, 1043
- Lorentz transformation**, 1035, 1035–1036
- Lorentz transformation equations**, 1035–1036
 consequences of, 1037–1038
 pairs of events, 1036t
- Loschmidt number**, 535
- L shell**, 1128–1129, 1128
- Lyman series**, 1100, 1101, 1104
- M**
- Mach cone**, 465, 465
- Mach cone angle**, 465, 465
- Mach number**, 465
- magic electron numbers**, 1185–1186
- magic nucleon numbers**, 1185–1186
- magnetically hard material**, 883
- magnetically soft material**, 883
- magnetic confinement**, 1212
- magnetic dipole moment**, 754, 754–755, 1114, 1114
- diamagnetic materials**, 875–876
- effective**, 1117
- ferromagnetic materials**, 875, 879
- orbital**, 873
- paramagnetic materials**, 875, 877–878
- spin**, 871–872, 872
- magnetic dipoles**, 738, 753–755, 754, 862, 862–863
- magnetic domains**, 879–880, 880
- and hysteresis, 880
- magnetic energy**, 811–812
- magnetic energy density**, 812–813
- magnetic field**, 736. *See also Earth's magnetic field*
- Ampere's law**, 771, 771–774
- circulating charged particle**, 744–747, 745, 746
- current-carrying wire**, 750, 750–751, 770, 770–771
- particle in magnetic field**, 736–738, 737
- magnetic flux**, 862
 and Faraday's law, 792–793
- magnetic flux linkage**, 805
- magnetic force**, 561, 736
- circulating charged particle**, 744–747, 745, 746
- current-carrying wire**, 750, 750–751, 770, 770–771
- particle in magnetic field**, 736–738, 737
- magnetic latitude**, 887
- magnetic levitation**, 697, 697
- magnetic materials**, 861, 862, 875
- diamagnetism**, 875, 875–876
- ferromagnetism**, 875, 879–880, 879
- magnetically hard and soft**, 883
- paramagnetism**, 875, 877, 877–878
- magnetic monopole**, 736, 862
- defined**, 735, 737, 739
- displacement current**, 866–868, 867
- due to current in circular arc of wire**, 767, 767–768
- due to current in long straight wire**, 765–767, 766
- Earth**, 870, 870–871
- energy density of**, 812–813
- energy stored in**, 811–812
- and Faraday's law of induction**, 792–794, 793
- Gauss' law for**, 862–863, 869t
- Hall effect**, 741, 741–744, 750
- induced**, 863–865, 864
- inside long straight current carrying wire**, 773, 773–774
- LC oscillations**, 826
- and Lenz' law**, 794, 794–797, 795
- outside long straight current carrying wire**, 772, 772–773
- parallel currents**, 770, 770–771
- selected objects and situations**, 738t
- solenoids**, 776–777, 775
- synchrotrons**, 749
- toroids**, 777, 777–778
- torque on current loops**, 752, 752–753
- traveling electromagnetic waves**, 890–893, 891, 893
- magnetic field lines**, 738, 738–739, 739
- current-carrying wire**, 766–767
- magnetic flux**, 862
 and Faraday's law, 792–793
- magnetic flux linkage**, 805
- magnetic force**, 561, 736
- circulating charged particle**, 744–747, 745, 746
- current-carrying wire**, 750, 750–751, 770, 770–771
- particle in magnetic field**, 736–738, 737
- magnetic latitude**, 887
- magnetic levitation**, 697, 697
- magnetic materials**, 861, 862, 875
- diamagnetism**, 875, 875–876
- ferromagnetism**, 875, 879–880, 879
- magnetically hard and soft**, 883
- paramagnetism**, 875, 877, 877–878
- magnetic monopole**, 736, 862
- magnetic potential energy**, 753–754, 811
- magnetic resonance**, 1120–1121, 1121
- magnetic resonance imaging**. *See MRI*
- magnetic wave component**, of electromagnetic waves, 892–893, 893
- magnetism**. *See also Earth's magnetic field*
 Earth, 870, 870–871
 electrons, 871–874
 protons and neutrons, 872
- magnetization**
 ferromagnetic materials, 879, 879–880
 paramagnetic materials, 877, 877–878
- magnetization curves**
 ferromagnetic materials, 879, 879–880
 hysteresis, 880, 880
 paramagnetic materials, 868, 877
- magnetometers**, 870
- magnets**, 561, 735, 735–739, 738, 869–871
 applications, 735
- magnetic dipoles**, 862, 862–863
- magnitude**
 of acceleration in one-dimensional motion, 19
 of acceleration in two- and three-dimensional motion, 62
- of displacement in one-dimensional motion**, 14
- of free-fall acceleration**, 25
- of vectors**, 39, 39
- of velocity in one-dimensional motion**, 15
- of velocity in two- and three-dimensional motion**, 62–63
- magnitude-angle notation**, vectors, 42
- magnitude ratio**, traveling electromagnetic waves, 893
- majority carriers**, 1153
 motions of, 1154–1156, 1155
- mantle (Earth)**, 353, 353
 density, 335
- mass**, 263t. *See also center of mass*
 defined, 90–91
 scalar nature of, 38, 91
 of selected objects, 7t

- mass (*cont.*)
 units, 6–7
 and wave speed on stretched string, 420
 weight contrasted, 96
- mass dampers, 394
- mass energy, 1043–1044
 energy equivalents of particles and objects, 1043t
- mass excess, 1171
- mass flow rate, 373
- massless, frictionless pulleys, 98, 99
 Newton’s laws applications, 100, 100–101
- massless cord, 98, 98
- massless spring, 149
- mass number, 572, 1168
 selected nuclides, 1168t
- mass spectrometer, 747, 747
- matter
 antimatter, 1222
 baryonic matter, 1240
 dark matter, 1240
 energy released by 1 kg, 1195t
 nuclear, 1174
 particle nature of, 1069–1070, 1069
 wave nature of, 1067–1069
- matter waves, 413
 electrons, 1067–1070, 1068, 1069, 1070, 1083, 1084
- maxima
 diffraction patterns, 990, 990
 double-slit interference, 964, 964, 967, 970
 single-slit diffraction, 992–995
 thin-film interference, 975
- Maxwell, James Clerk, 516, 523, 561, 771, 863, 890, 899, 958, 1236
- Maxwellian electromagnetism, 1218
- Maxwell’s equations, 861, 869t, 1071
- Maxwell’s law of induction, 863
- Maxwell’s rainbow, 889–890
- Maxwell’s speed distribution law 516–519, 517
- mean free time, 694
- mean life, radioactive decay, 1175, 1219
- measurement, 1
- mechanical energy
 conservation of, 173–175
 and conservation of total energy, 183
- damped harmonic oscillator, 400
 satellites in orbit, 345, 345
 simple harmonic motion, 392–393, 393
- mechanical waves, 413
- medium, 893
- melting point, 486
 selected substances, 487t
- meniscus concave lens, 957
- meniscus convex lens, 957
- mercury barometer, 361, 365, 365
- mercury thermometer, 482
- mesons, 1221, 1228–1229
 and eightfold way, 1231t, 1231–1232
 and quark model, 1232–1234, 1237
 underlying structure suggested, 1232
- messenger particles, 1235–1237
- metallic conductors, 682, 696
- metal-oxide-semiconductor-field effect-transistor (MOSFET), 1159, 1159
- metals
 coefficients of linear expansion, 482t
 density of occupied states, 1149, 1149–1150
 density of states, 1146–1148, 1147
 elastic properties of selected, 317t
 electrical properties, 1145–1150
 lattice, 315, 315
 occupancy probability, 1147–1148, 1148
 resistivities of selected, 689t
 speed of sound in, 447t
 thermal conductivity of selected, 494t
 unit cell, 1142
- metastable states, 1132
- meter (m), 1, 3–4
- metric system, 2
- mho (unit), 690
- Michelson’s interferometer, 980, 980
- microfarad, 657
- microscopes, 943
 compound, 944, 944
 electron microscopes, 1080
- microscopic clocks, time dilation tests, 1029
- microstates, in statistical mechanics, 550 – 552
- microwave cooking, 595
- microwaves, 413, 889
- Doppler effect, 461
- Mid-Atlantic Ridge, magnetism, 870, 870
- Milky Way galaxy, 330
- Millikan oil-drop experiment, 592, 592
- millimeter of mercury (mm Hg), 361
- millimeter-wave radar, 1017
- miniature black holes, 351
- minima
 circular aperture diffraction, 1000, 1000
 diffraction patterns, 990, 990–991, 991
 double-slit interference, 964, 964, 966, 970
 single-slit diffraction, 992–995, 995–997, 996
 thin-film interference, 976
- minority carriers, 1153
 motions of, 1154–1155
- mirage, 925, 925
- mirrors, 926
 plane, 926, 926–928
 spherical, 928–934, 928, 929, 946, 946–947
- mobile radio, 889
- moderators, for nuclear reactors, 1202
- modulus of elasticity, 315
- Mohole, 353
- molar mass, 508
- molar specific heat, 485
 at constant pressure, 521–522
 at constant volume, 519–520
 and degrees of freedom, 523, 523–524
 selected materials, 485t
- molecular mass, 508
- molecular speeds, Maxwell’s distribution of, 516–519, 517
- molecules, 1112
- matter wave interference, 1069
- moment arm, 259, 259
- moment of inertia, 253
- momentum, 210. *See also* angular momentum; linear momentum
 center of momentum frame, 1056
 and kinetic energy, 1046
 and relativity, 1042–1043
- monatomic molecules, 519
 degrees of freedom, 523, 523–524, 524t
- molar specific heats at constant volume, 520t
- monochromatic light, 907
- lasers, 1131
- monovalent atom, 1145
- Moon, 330
 corona around, 1018, 1018
 potential effect on humans, 351
- radioactive dating of rocks, 1183
- more capacitive than inductive circuit, 844
- more inductive than capacitive circuit, 844
- Moseley plot, 1129, 1129–1130
- MOSFET (metal-oxidesemiconductor-field-effect-transistor), 1159, 1159
- most probable configuration, 552
- most probable speed in fusion, 1207
 of gas molecules, 517
- motion, 13. *See also* acceleration; collisions; displacement; forces; one-dimensional motion; position; uniform circular motion; velocity
 graphical integration, 27, 27–28
- projectile, 64–70, 65
- relative in one dimension, 73, 73–74
- relative in two dimensions, 74–75, 75
- two- and three-dimensional, 58–64
- MRI (magnetic resonance imaging)
 effect on magnetic tattoo inks, 861, 861
- M shell, 1128, 1128
- multiloop circuits, 707, 714, 714–720
- resistances in parallel, 714, 714–720
- multimeter, 720
- multiplication, of vectors, 47–52, 48
- multiplication factor, nuclear reactors, 1203
- multiplicity, of configurations in statistical mechanics, 551
- muon neutrinos, 1227, 1227t
- muons, 1029–1030, 1054, 1219, 1227t, 1227
- decay, 1225
 from proton-antiproton annihilation, 1224t, 1224
- musical sounds, 456–459
- mutual induction, 813–815, 814

- N**
- nanotechnology, 1083
 - National Institute of Standards and Technology (NIST), 5
 - natural angular frequency, 402, 836
 - series *RLC* circuits, 846, 846
 - NAVSTAR satellites, 1022
 - n* channel, in MOSFET, 1159
 - near point, 943, 943
 - negative charge, 562, 562
 - negative charge carriers, 684
 - drift speed, 686
 - negative direction, 13, 13
 - negative lift, in race cars, 126–127, 127
 - negative terminal, batteries, 658, 658, 706
 - net force, 92
 - net torque, 259
 - net wave, 425
 - net work, 144
 - net work per cycle
 - Carnot engine, 544
 - neutral equilibrium, potential energy curves, 178
 - neutral pion, 1023
 - neutrinos, 1180
 - and beta decay, 1180–1181
 - in fusion, 1209–1210
 - as leptons, 1221
 - as nonbaryonic dark matter, 1240
 - from proton-antiproton annihilation, 1224t
 - neutron diffraction, 1069
 - neutron excess, 1169
 - neutron generation time, in nuclear reactors, 1215
 - neutron number, 1168
 - selected nuclides, 1168t
 - neutron rich fragments, 1197
 - neutrons, 563, 1219
 - accelerator studies, 748
 - balance in nuclear reactors, 1202–1203, 1203
 - charge, 571, 571t
 - collision with protons, 1190
 - control in nuclear reactors, 1202–1205, 1203
 - discovery of, 1236
 - and eightfold way, 1231t
 - as fermions, 1220
 - formation in early universe, 1241
 - as hadrons, 1221
 - magnetic dipole moment, 754
 - and mass number, 572
 - as matter wave, 1068
 - spin angular momentum, 871
 - neutron stars, 82, 352, 1190
 - density of core, 360t
 - escape speed, 341t
 - magnetic field at surface, 738t
 - newton (N), 88
 - Newton, Isaac, 87, 330, 342, 991
 - Newtonian form, of thin-lens formula, 956
 - Newtonian mechanics, 87, 1071, 1218
 - Newtonian physics, 1083
 - newton per coloumb, 580
 - Newton's first law, 87–91
 - Newton's law of gravitation, 330–333, 331
 - Kepler's laws follow from, 342
 - Newton's laws, 87
 - applying in problem solving, 100–105
 - Newton's rings, 987, 987
 - Newton's second law, 91–94
 - angular form, 285–286
 - applying in problem solving, 100–105
 - for rotation, 260–262
 - system of particles, 206–209, 207
 - in terms of momentum, 210
 - translational and rotational forms contrasted, 263t, 289t
 - units in, 92t
 - Newton's third law, 98
 - NMR spectrum, 1121, 1121
 - noble gases, 1126, 1186
 - nodes, 431, 431, 433
 - nonbaryonic dark matter, 1240
 - nonconducting sheet
 - Gauss' law, 617, 617
 - nonconductors, 563
 - nonconservative forces, 168
 - noninertial frame, 89
 - nonlaminar flow, 371
 - nonpolar dielectrics, 672
 - nonpolar molecules, 638
 - nonquantized portion, of energy level diagram, 1092, 1092
 - nonsteady flow, 371
 - nonuniform displacement current, 885
 - nonuniform displacement-current density, 885
 - nonuniform electric field, 608, 884
 - nonuniform electric flux, 884
 - nonuniform magnetic field

- diamagnetic materials, 875
- and loop model for electrons, 873–874, 873, 874
- nonviscous flow, 371
- normal (optics), 904–906, 905, 906
- normal force, 96–98, 97
- normalizing, wave function, 1089
- north magnetic pole, 870
- north pole, magnets, 738–739, 739, 862, 862
- n*-type semiconductors, 1152, 1152–1153. *See also p-n junctions*
- nuclear angular momentum, 1172
- nuclear binding energy, 1171–1172, 1172, 1198, 1199
 - and excitation energy, 1110
 - selected nuclides, 1168t
- nuclear energy, 1195
- nuclear energy levels, 1172
- nuclear fission, 1172, 1196–1201, 1199
- nuclear force, 1173–1174
- nuclear fusion, *See thermonuclear fusion*
- nuclear magnetic moment, 1172
- nuclear magnetic resonance (NMR), 1121
- nuclear magnetic resonance (NMR) spectrum, 1121, 1121
- nuclear power plant, 1202, 1204
- nuclear radii, 1170
- nuclear reactions, 1043
- nuclear reactors, 1202–1205
 - natural, 1206, 1206
- nuclear spin, 1172
- nuclear weapons, 1172
- nucleons, 1168, 1221
 - magic nucleon numbers, 1186
- nucleus, 563, 1165–1168
 - models, 1184–1186, 1185
 - radioactive decay, 572, 1219–1220
- nuclides, 1167–1168, 1168t. *See also radioactive decay*
 - magic nucleon numbers, 1185–1186
 - organizing, 1169–1170, 1169
- nuclidic chart, 1169–1170, 1169
 - and radioactivity, 1181–1182, 1182
- number density of charge carriers, 742, 1143
- semiconductors, 1150–1152
- silicon and copper compared, 1143t
- number density of conduction electrons, 1145

O

- object, 926, 926–927
- object distance, 926
- objective
 - compound microscope, 944, 944
 - refracting telescope, 945, 945
- occupancy probability, 1147–1148, 1148
- occupied levels, 1122
- occupied state density, 1149, 1149–1150
- ohm (unit), 689
- ohmic losses, 850
- ohmmeter, 689
- ohm-meter, 689
- Ohm's law, 692, 692
 - microscopic view of, 693, 693–694
- oil slick, interference patterns from, 973
- one-dimensional elastic collisions, 221, 221–223
- one-dimensional electron traps multiple electrons, 1122
 - single electron, 1084–1096
- one-dimensional explosions, 216, 216
- one-dimensional inelastic collisions, 218, 218–221, 219
- one-dimensional infinite potential well, 1085
- one-dimensional motion, 13
 - acceleration, 18–27
 - average acceleration, 18
 - average speed, 16
 - average velocity, 14–16
 - constant acceleration, 22, 22–25
 - displacement, 13–14
 - free-fall acceleration, 25–26
 - graphical integration, 27, 27–28
 - instantaneous acceleration, 18
 - instantaneous velocity, 17–18
 - position, 13, 13–14
 - relative, 73, 73–74
 - Schrödinger's equation, 1071
 - speed, 18
 - velocity, 17–18

one-dimensional variable force, 151–152, 152
 one-half rule, for intensity of transmitted polarized light, 902
 one-way processes, 536
 open-tube manometer, 365, 365–366
 optical fibers, 912, 1131
 with junction lasers, 1158
 optical instruments, 943–946
 optical interference, 959. *See also* interference and rainbows, 962, 962
 optics, 889
 orbital angular momentum, 872, 873, 1115–1117, 1116
 and orbital quantum number, 1115t
 orbital energy, 1098
 orbital magnetic dipole moment, 872, 1115–1117, 1116
 diamagnetic materials, 875–876
 ferromagnetic materials, 875, 879
 paramagnetic materials, 877–878
 orbital magnetic quantum number, 873, 1101t, 1102, 1115t
 orbital quantum number, 1101t, 1102, 1115t, 1143
 orbital radius, 1097–1098
 order numbers, diffraction gratings, 1006
 order of magnitude, 4
 origin, 13
 oscillating circuit, 826
 oscillation mode, 433–434
 oscillations, 386. *See also* pendulums; simple harmonic motion; waves forced, 402, 836–837, 837 free, 402
 LC circuits, 826–829, 827
 out of phase
 ac circuits, 842t
 capacitive load, 838–839
 inductive load, 840
 simple harmonic motion, 390
 sound waves, 451
 thin-film interference, 975
 waves, 425
 overpressure, 366
 oxygen, 523
 distribution of molecular speeds at 300 K, 517

heats of transformation, 487t
 molar specific heat and degrees of freedom, 524t
 molar specific heat at constant volume, 520t
 paramagnetism of liquid, 877
 RMS speed at room temperature, 512t

P

pair production, 572–573
 pancake collapse, of tall building, 237
 parallel-axis theorem, for calculating rotational inertia, 254–255, 255
 parallel circuits
 capacitors, 662–663, 663, 715t
 inductors, 822
 resistors, 714, 714–717, 715t
 summary of relations, 715t
 parallel components, of unpolarized light, 912
 parallel currents, magnetic field between two, 770, 770–771
 parallel-plate capacitors, 656, 656
 calculating capacitance, 659
 with dielectric, 671, 671–674, 672
 displacement current, 866–868, 867
 energy density, 668
 induced magnetic field, 863–866
 paramagnetic material, 875, 875, 877–878
 paramagnetism, 875, 877, 877–878
 parent nucleus, 572
 partial derivatives, 419, 450, 895
 partially occupied levels, 1122
 partially polarized light, 901, 902
 particle accelerators, 747–749, 1218, 1219, 1219
 particle-antiparticle annihilation, 1222
 particle detectors, 1219, 1219
 particles, 13, 571
 antiparticles, 1222, 1224
 basic forces and messenger particles, 1235–1237
 bosons, 1220–1221, 1221
 circulating charged, 744–747, 745, 746

conservation of baryon number, 1229
 conservation of lepton number, 1228
 conservation of strangeness, 1231
 decay processes, 1222–1227
 fermions, 1220–1221, 1221
 formation in early universe, 1241
 hadrons, 1221, 1228–1229
 leptons, 1221, 1227t, 1227–1228
 magnetic force on, 736–738, 737
 particle nature of matter, 1069–1070, 1069
 quark model, 1232–1235, 1233
 Standard Model, 1220–1222
 particle systems
 angular momentum, 288
 center of mass, 201–206, 202, 205
 linear momentum, 211
 Newton's second law for, 206–209, 207
 pascal (Pa), 361, 446, 900
 Pascal's principle, 366, 366–367
 Paschen series, 1100, 1101
 path-dependent quantities, 490
 path-independent quantities
 conservative forces, 168–170, 169
 gravitational potential energy, 340–341
 work done by electrostatic force, 628
 path length difference
 double-slit interference, 965, 965–966, 971–972
 single-slit diffraction, 993, 993, 994, 995
 sound waves, 451
 thin-film interference, 973
 Pauli exclusion principle, 1121
 bosons don't obey, 1220
 and energy levels in crystalline solids, 1143
 fermions obey, 1220
 and Fermi speed, 1145
 nucleons, 1185
 and periodic table, 1125
 pendulums
 as angular simple harmonic oscillator, 394, 394–395
 ballistic, 220, 220
 conservation of mechanical energy, 173–175, 174
 physical, 396–397, 397
 simple, 395, 395–396

torsion, 394, 394
 underwater swinging (damped), 400
 perfect engines, 546, 546
 perfect refrigerators, 549, 549
 perihelion distance, 344
 period
 circulating charged particle, 744–747
 LC oscillations, 826
 linear simple harmonic oscillator, 390
 simple harmonic motion, 387, 388, 389
 sound waves, 449
 waves, 416, 416–417
 periodic motion, 386
 periodic table, 1057, 1113
 building, 1124–1127
 x rays and ordering of elements, 1127–1130
 period of revolution, 71
 permanent electric dipole moment, 638
 permanent magnets, 736
 permeability constant, 765
 permittivity constant, 566
 perpendicular components, of unpolarized light, 912
 phase
 alternating current, 842t
 simple harmonic motion, 388, 388
 waves, 415, 415
 phase angle
 alternating current, 842t
 simple harmonic motion, 388, 388
 phase change, 486
 phase constant
 alternating current, 842t, 848
 series *RLC* circuits, 844–845, 845
 simple harmonic motion, 388, 388
 waves, 417, 417
 phase difference
 double-slit interference, 965, 969–970, 971–972
 Michelson's interferometer, 980
 optical interference, 961–962
 and resulting interference type, 427t
 single-slit diffraction, 995
 sound waves, 451
 thin-film interference, 975
 waves, 425–426
 phase-shifted
 reflection phase shifts, 974, 974

- sound waves, 451
 waves, 425–426
 phasor diagram, 428–429, 429
 phasors, 428–430, 430
 capacitive load, 838, 838, 839
 double-slit interference, 971
 inductive load, 840, 840, 841
 resistive load, 838
 series *RLC* circuits, 844–845,
 845
 single-slit diffraction,
 995–999, 996, 998
 photodiode, 1158
 photoelectric current, 1059
 photoelectric effect,
 1059–1062, 1061
 photoelectric equation,
 1060–1061
 photoelectrons, 1059
 photomultiplier tube, 1067
 photon absorption, 1058, 1113
 absorption lines, 1100, 1101
 for electrons to gain energy,
 1086
 energy changes in hydrogen
 atom, 1098
 lasers, 1132
 photon emission, 1058, 1113
 for electrons to lose energy,
 1086
 emission lines, 1100, 1101
 energy changes in hydrogen
 atom, 1098
 lasers, 1132, 1132–1133
 stimulated emission, 1132,
 1133
 photons
 as bosons, 1220
 in early universe, 1241
 energy of, 1058
 momentum, 1062,
 1062–1065, 1063
 and photoelectric effect,
 1059–1062
 as quantum of light,
 1057–1059
 single-photon version of
 double-slit interference,
 1065, 1065–1066
 single-photon wide-angle
 version of double-slit
 interference, 1067, 1067
 virtual, 1235–1236
 physical pendulum, 396–398,
 398
 picofarad, 657
 piezoelectricity, 1075
 pinhole diffraction, 990
 pions, 1024, 1055, 1219
 decay, 1225
 and eightfold way, 1231t
 as hadrons, 1221
 as mesons, 1221
 from proton-antiproton an-
 nihilation, 1224t,
 1222–1225
 pitch, 359
 pitot tube, 383
 planar symmetry, Gauss' law,
 617, 617–618
 planar waves, 446
 Planck constant, 1058
 plane-concave lens, 957
 plane-convex lens, 957
 plane mirrors, 926–928, 927
 plane of incidence, 906
 plane of oscillation, polarized
 light, 901, 901
 plane of symmetry, center of
 mass of solid bodies
 with, 203
 plane-polarized waves, 901
 plane waves, 891
 plastics
 as insulators, 563
 rubbing rod with fur, 563
 plates, capacitors, 657, 657–658
 plate tectonics, 13
 plum pudding model, of atom,
 1166
p-n junction diode, 692, 697
p-n junctions, 1154–1156, 1155
 junction lasers, 1158, 1158
 junction rectifiers, 1156,
 1156–1157, 1164
 light-emitting diodes
 (LEDs), 1157–1158,
 1158
 point charges
 Coulomb's law, 565, 565–572
 in electric field, 592–593
 electric field due to, 582,
 582–583
 electric potential due to, 635,
 635
 electric potential due to
 group of, 636, 636–637
 electric potential energy of
 system, 642, 642–644
 point image, 926–927
 point of symmetry, center of
 mass of solid bodies
 with, 203
 point source, 446
 isotropic, 454, 898
 light, 898, 926
 polar dielectrics, 671
 polarity
 of applied potential differ-
 ence, 692
 of Earth's magnetic field, re-
 versals in, 870, 870
 polarization, 901, 901–904, 902,
 903
 by reflection, 912, 912–913
 polarized light, 901, 901–902
 intensity of transmitted, 902,
 902–904, 903
 polarized waves, 901, 901–904,
 902
 polarizer, 903
 polarizing direction, 903, 903
 polarizing sheets, 902, 902–903,
 903, 912
 polar molecules, 638
 Polaroid filters, 902
 pole faces, horseshoe magnet,
 739
 polyatomic molecules, 520
 degrees of freedom, 523,
 523–524, 524t
 molar specific heats at con-
 stant volume, 520t
 polycrystalline solids, 880
 population inversion, 1133,
 1139
 junction lasers, 1158
 porcelain
 dielectric properties, 69t,
 671
 position, 263t. *See also* angular
 position; displacement
 one-dimensional motion, 13,
 13–14
 relating linear to angular,
 250
 simple harmonic motion,
 388
 two- and three-dimensional
 motion, 58–59, 59
 position vector, 58, 58
 positive charge, 672, 562
 positive charge carriers, 684
 drift speed, 686
 emf devices, 706
 positive direction, 13, 13
 positive ions, 563
 positive kaons, 1030
 positive terminal, batteries,
 658, 658, 706
 positrons
 antihydrogen, 1222
 bubble chamber tracks, 572,
 737
 electron-positron annihila-
 tion, 572, 572, 1222
 in fusion, 1207–1208
 potential, 629–631. *See also*
 electric potential
 potential barrier, 1074, 1074,
 1200
 potential difference, 629–631
 across real battery, 712–713
 capacitors, 659, 659–662
 capacitors in parallel,
 662–663, 663
 capacitors in series, 664,
 663–666, 842–844
 Hall, 741
LC oscillations, 826
 and Ohm's law, 692
 power, potential, and emf,
 712–713
 resistors in parallel, 714–715
 resistors in series, 710, 710,
 842–844
RL circuits, 807, 807–810
 single-loop circuits, 707,
 707–709
 between two points in cir-
 cuit, 711, 711–713, 712
 potential energy. *See also* elas-
 tic potential energy;
 gravitational potential
 energy
 and conservation of me-
 chanical energy, 173,
 173–175, 174
 and conservation of total en-
 ergy, 183–184
 defined, 167
 determining, 170–172
 electric dipole, 594–595
 satellites in orbit, 345, 345
 simple harmonic motion,
 392, 393
 and work, 167, 167–170
 yo-yo, 280–281
 potential energy barrier, 1074,
 1074
 potential energy curves,
 176–179, 177
 potential energy function,
 176–178, 177
 potential method, of calculat-
 ing current in single-
 loop circuits, 707–708
 potential well, 178
 potentiometer, 670
 pounds per square inch (psi),
 361
 power, 155–156, 156, 185, 263t
 in alternating current cir-
 cuits, 847–848
 in direct current circuits,
 695–696
 and emf in circuits, 712–713
 in rotation, 262
 traveling wave on stretched
 string, 421, 421–423
 power factor, 849
 power lines
 transformers for, 850

power transmission systems, 682
transformer requirements, 850
Poynting vector, 897–899, 898
precession, of gyroscope, 293–295, 295
precession rate, of gyroscope, 294
prefixes, for SI Units, 2t
pressure
 fluids, 359–361
 and ideal gas law, 508–511
 measuring, 365, 365–366
 and RMS speed of ideal gas, 511–513
 scalar nature of, 38
 as state property, 537
 triple point of water, 478
 work done by ideal gas at constant, 510
pressure amplitude
 sound waves, 449, 449
pressure field, 580
pressure sensor, 360
pressurized-water nuclear reactor, 1204, 1204
primary coil, transformer, 850
primary loop, pressurized-water reactor, 1204, 1204
primary rainbow, 909, 922
primary rainbows and optical interference, 962, 962
primary winding, transformer, 850
principal quantum number, 1102, 1101t, 1115t, 1143
principle of conservation of mechanical energy, 173
principle of energy conservation, 140
principle of equivalence, 347
principle of superposition, 88
 and electrostatic force, 566
 for gravitation, 333–334
 for waves, 425, 425
prism, 908, 908
 angle of minimum deviation, 920
probability
 and entropy, 552–553
probability density, 1071–1072
 barrier tunneling, 1074–1075
 finding, 1072
 trapped electrons, 1088–1090, 1089
probability distribution function, 516
probability of detection
 hydrogen electron, 1102–1103

trapped electrons, 1088–1089
probability wave
 light as, 1066–1067
 matter wave as, 1068
projectile, 64
 elastic collisions in one dimension with moving target, 222–223
 elastic collisions in one dimension with stationary target, 221–222
 inelastic collisions in one dimension, 218
 series of collisions, 213
 single collision, 211–212
projectile motion, 64–70, 65, 66
proper frequency, 1040
proper length, 1032
proper period, 1042
proper time, 1028
proper wavelength, 1041
proton number, 1168
 selected nuclides, 1168t
proton-proton (p-p) cycle, 1209–1210, 1209
protons, 563, 1219
 accelerator studies, 748
 and atomic number, 572
 as baryons, 1221
 charge, 571, 571t
 decay rate, 1245
 in equilibrium, 569
 as fermions, 1220
 formation in early universe, 1241
 in fusion, 1207–1213
 as hadrons, 1221
 magnetic dipole moment, 754, 754t
 mass energy, 1043t
 and mass number, 572
 as matter wave, 1068, 1069, 1083
 spin angular momentum, 871
 ultrarelativistic, 1047
 in well, 652
proton synchrotrons, 749
p subshells, 1125, 1126
p-type semiconductors, 1153, 1153
pulleys, massless frictionless, 98, 98
 Newton's laws applications, 100, 100–101
pulsars, 268
 secondary time standard based on, 9
pulse, wave, 414, 414

Q
quadrupole moment, 600
quanta, 1057
quantization, 1057, 1083
 electric charge, 571
 energy in electron traps, 1084–1088
 orbital angular momentum, 872
 orbital energy, 1098
 spin angular momentum, 871
 quantum, 1057
quantum chromodynamics (QCD), 1236
quantum corrals, 1094, 1094
quantum dots, 1083, 1094, 1094
quantum electrodynamics (QED), 872, 1235
quantum jump, 1086
quantum mechanics, 87, 1057
quantum numbers, 1084, 1115t
 charge, 1224
 for hydrogen, 1102–1103
 and Pauli exclusion principle, 1121
 and periodic table, 1124–1126
quantum physics, 1057. *See also* electron traps; Pauli exclusion principle; photons; Schrödinger's equation
barrier tunneling, 1074, 1074–1076
and basic properties of atoms, 1112–1114, 1113
confinement principle, 1084
correspondence principle, 1089, 1106
Heisenberg's uncertainty principle, 1072, 1073, 1236
hydrogen wave function, 1099–1106
matter waves, 1083
nucleus, 1165
occupancy probability, 1147–1148, 1148
particles, 1219
and solid-state electronic devices, 1142
quantum states, 1083, 1114
 degenerate, 1095
 density of, 1146–1147, 1147
 density of occupied, 1149, 1149–1150
 hydrogen with $n = 2$, 1105, 1105–1106
quantum theory, 523, 525, 1057, 1083

quantum transition, 1086
quantum tunneling, 1074, 1074–1076
quark family, 1233t
quark flavors, 1232, 1236–1237
quarks, 578, 654, 747, 1232–1235, 1233, 1233t
 charge, 571
 formation in early universe, 1241
quasars, 1053, 1237
 gravitational lensing associated with, 348
quicksand, 385

R
R-value, 494
race cars
 collision with wall, 214, 214
 negative lift in Grand Prix cars, 126–127, 127
 pit stop fuel dispenser fire, 723, 723–724
rad (unit), 1184
radar waves, 413
radial component
 of linear acceleration, 251
 of torque, 259
radial probability density, 1103, 1103, 1106
radians, 43, 242
radiated waves, 890
radiation (electromagnetic), 890
radiation dosage, 1184
radiation heat transfer, 496–497
radiation pressure, 899–900
radiation therapy, 1190
radioactive dating, 1183, 1183
radioactive decay, 572, 1169, 1219–1220
 alpha decay, 1177–1179, 1178
 beta decay, 1180–1181, 1181, 1234
 muons, 1029
 and nuclidic chart, 1181–1182, 1182
 process, 1174–1177
radioactive elements, 1165
radioactive wastes, 1204, 1204
radionuclides, 1169
radiation dosage, 1184
radio waves, 413, 889, 890
 Doppler effect, 461
radium, 1191
radius of curvature
 spherical mirrors, 928, 928–929, 929
spherical refracting surfaces, 934, 934

- radon, 1165, 1191
 rail gun, 770, 770–771, 790, 790
 rainbows, 888, 908–909, 909,
 922
 and optical interference,
 962, 962
 ramp
 rolling down, 278–279, 279
 randomly polarized light,
 901, 901–902
 range, in projectile motion, 64,
 68
 rare earth elements, 1129
 paramagnetism, 875
 rattlesnake, thermal radiation
 sensors, 496
 ray diagram, 932, 932
 Rayleigh's criterion, 1000,
 1000–1001
 rays, 446, 446
 locating extended object im-
 ages by drawing, 939,
 939
 locating images by drawing,
 932, 932
 RBE (relative biology effec-
 tiveness) factor, 1184
RC circuits, 720, 720
 capacitor charging, 720,
 720–721
 capacitor discharging, 720,
 720, 722–724
 real batteries, 707, 707
 internal resistance, 711, 711
 potential difference across,
 711, 711–712
 real emf devices, 707, 707
 internal resistance, 711
 real engines
 efficiency, 546, 549–550
 real fluids, 371
 real focal point, 929, 929
 real images, 924–925
 spherical mirrors, 930
 spherical refracting surfaces,
 934, 934
 thin lenses, 938, 938
 real solenoid, 776
 recessional speed parameter,
 Doppler shift, 1246
 rechargeable batteries, 707
 recharging batteries, 712
 red shift, 1041
 reference circle, 399
 reference configuration, for
 potential energy, 171,
 628
 reference frames, 73–74
 inertial, 89–90
 noninertial, 89
 reference line, 242, 242
 reference particle, 399
 reference point, for potential
 energy, 171
 reference potential energy, 628
 reflected light, 904
 reflected radiation, 899
 reflected ray, 904–905, 905
 reflecting planes, 1012, 1012
 reflection, 904–907, 905, 906
 law of, 906
 polarization by, 912, 912–913
 of standing waves at bound-
 ary, 432, 432
 total internal, 911, 911–912
 reflection phase shifts, 974,
 974–976
 refracted light, 906
 refracted ray, 906, 906
 refracting telescope, 945,
 945–946
 refraction, 904–910, 906. *See*
 also index of refraction
 and chromatic dispersion,
 907, 907–908, 908
 law of, 906, 959–962, 959
 refrigerators, 548, 548–549
 relative biology effectiveness
 (RBE) factor, 1184
 relative motion
 in one dimension, 73–74, 73
 in two dimensions, 74–75, 74
 relativity, 1057, 1218
 brief description, 1023
 Doppler effect for light,
 1040–1042, 1041
 and energy, 1043–1047
 general theory of, 347–348,
 1022–1023, 1029
 and kinetic energy, 1045,
 1045–1047
 of length, 1031, 1031–1034,
 1037
 Lorentz factor, 1028, 1028,
 1029, 1043
 Lorentz transformation,
 1035, 1035–1036
 and mass energy, 1043–1045,
 1043t
 measuring events,
 1024–1025, 1025
 and momentum, 1042
 postulates, 1023
 of simultaneity, 1025–1026,
 1026, 1037
 special theory of, 87, 893,
 1022, 1029
 speed of light postulate,
 1023–1024
 of time, 1027–1031, 1027,
 1037, 1042
 of total energy, 1044–1045
 of velocities, 1039, 1039
 relativity postulate, 1023
 relaxed state, of spring, 149,
 149
 released energy, from fusion
 reaction, 1045
 rem (unit), 1184
 resistance, 689–691
 alternating current, 842t
 calculating from resistivity,
 690, 690
 Ohm's law, 692, 692–694
 parallel circuits, 714,
 714–719
RC circuits, 720–723
RL circuits, 807–810
RLC circuits, 833–834, 842–844
 series circuits, 710, 710,
 842–844
 resistance rule, 709
 resistivity, 689, 1142
 calculating resistance from,
 690, 690
 Ohm's law, 693
 selected materials at room
 temperature, 689t
 semiconductors, 1151
 silicon and copper com-
 pared, 696t, 696–697,
 1143t
 resistors, 689, 689–690
 with ac generator, 836,
 836–837
 in multiloop circuits, 714,
 714–719, 717
 Ohm's law, 692, 692–694
 in parallel, 714, 714–719
 phase and amplitude rela-
 tionships for ac circuits,
 842t
 power dissipation in ac cir-
 cuits, 847–848
 and power in circuits,
 695–696
RC circuits, 720, 720–723
RL circuits, 807, 807–810
RLC circuits, 833, 833–834,
 842–844, 843
 in series, 710, 710, 842–844,
 843
 series *RLC* circuits, 842–844,
 843
 in single-loop circuits, 707,
 707–710
 work, energy, and emf, 706,
 706–707
 resolvability, 1000, 1000–1001
 resolving power
 diffraction gratings,
 1009–1011, 1010
 highest for microscope, 1080
 refracting telescope, 945, 945
 resolving vectors, 41
 resonance
 forced oscillations, 402, 837
 series *RLC* circuits, 845–846,
 846
 and standing waves, 433,
 433–435
 resonance capture, of neu-
 trons in nuclear reac-
 tors, 1202
 resonance condition cy-
 clotrons, 748
 resonance curves
 series *RLC* circuits, 845–846,
 846
 resonance peak, 402, 1121
 resonant frequencies, 433, 433,
 434
 musical sounds, 457, 458
 response time, nuclear reactor
 control rods, 1203
 rest energy, 1043
 rest frame, 1029
 rest length, 1032
 restoring force, 149
 resultant, of vector addition, 39
 resultant force, 92
 resultant torque, 259
 resultant wave, 425, 425
 reverse saturation current,
 junction rectifiers, 1164
 reversible processes, 537–541
 right-handed coordinate sys-
 tem, 44, 44
 right-hand rule, 247, 247
 Ampere's law, 771, 771
 angular quantities, 247, 247
 displacement current,
 866–867, 867
 induced current, 795, 795
 Lenz's law, 794, 794
 magnetic dipole moment,
 753–754, 754
 magnetic field due to cur-
 rent, 765, 765
 magnetic force, 737, 737
 magnetism, 770
 vector products, 50, 50, 769
 rigid bodies
 angular momentum of rota-
 tion about fixed axis,
 288–290, 289
 defined, 241
 elasticity of real, 315
 ring charge distributions, 586,
 586–587, 590
RL circuits, 807, 807–810, 809
RLC circuits, 833, 833–834
 series, 842–846, 843

RMS (root-mean-square) current, in ac circuits, 847
 RMS (root-mean-square) speed, of ideal gas, 511, 511–513, 516
 selected substances, 512t
 rock climbing
 crimp hold, 324, 324
 energy conservation in descent using rings, 183, 183
 energy expended against gravitational force
 climbing Mount Everest, 199
 friction coefficients between shoes and rock, 119
 lie-back climb along fissure, 322, 322
 rockets, 224–226, 225
 roller-coaster headache, 252
 roller coasters
 maximum acceleration, 20
 rolling
 down ramp, 278, 278–280
 forces of, 278–280, 279
 friction during, 278, 278
 kinetic energy of, 277, 277
 as pure rotation, 276–277, 276
 as translation and rotation combined, 275–277, 277
 yo-yo, 280–281, 281
 room temperature, 476
 root-mean-square (RMS)
 current, in ac circuits, 847
 root-mean-square (RMS) speed, of ideal gas, 511, 511–513
 and distribution of molecular speeds, 516
 selected substances, 512t
 rotation, 242. *See also* rolling; torque; terms beginning with “angular”
 angular momentum of rigid body rotating about fixed axis, 288–290, 289
 conservation of angular momentum, 290–293, 291, 293
 constant angular acceleration, 248–249
 kinetic energy of, 253–254, 254
 Newton’s second law for, 260–261
 relating linear and angular variables, 250–253, 251
 in rolling, 275–277, 276
 rotational equilibrium, 307

rotational inertia, 253, 263t
 calculating, 254–258
 rotational kinetic energy, 253
 of rolling, 278
 and work, 262–264
 yo-yo, 280–281
 rotational symmetry, 582
 rotational variables, 241–246, 289t
 vector nature of, 246–247
 rotation axis, 242, 242
 Rotor (amusement park ride), 249
 Rowland ring, 879, 879
 rubber band, entropy change on stretching, 542
 rulers, 1
 rulings, diffraction grating, 1003, 1003
 Rutherford, Ernest, 1081, 1165
R-value, 494
 Rydberg constant, 1099

S

satellites
 charge buildup in radiation belt, 623
 geosynchronous orbit, 354
 gravitational potential energy, 339
 Kepler’s laws, 342–344
 orbits and energy, 345, 345–346
 saturating magnetization, 877, 878, 879, 887
 scalar components, 44
 scalar fields, 580
 scalar product, 48
 scalars, 38
 multiplying vectors by, 47–48
 scanning tunneling microscope (STM), 1075–1076, 1075
 quantum corral, 1094, 1094
 scattering
 polarized light, 903
 x rays, 1011, 1011
 schematic diagrams, 658
 Schrödinger’s equation, 1071–1072, 1083
 electron in finite well, 1091–1092
 electron in one-dimensional infinite well, 1088–1091
 electron in rectangular box, 1095
 electron in rectangular corral, 1095
 finding probability density, 1072
 and hydrogen, 1099–1106

hydrogen ground state, 1102–1103, 1103, 1115
 numerical solution for multicomponent atoms, 1124
 scientific notation, 2
 scintillation counter, 1248
 seat of emf, 705
 secondary coil, transformer, 850–851
 secondary loop, pressurized water reactor, 1204, 1204
 secondary maxima, diffraction patterns, 990, 990
 secondary rainbows, 909, 909, 922, 962
 secondary standards, 3
 secondary winding, transformer, 850–851
 second law of thermodynamics, 541–542
 second minima, interference patterns, 967
 second-order bright fringes, interference patterns, 966
 second-order dark fringes, interference patterns, 967
 second-order line, 1006
 second reflection point, 921, 921
 second side maxima, interference patterns, 966
 secular equilibrium, 1191
 seismic waves, 413
 self-induced emf, 806, 806
 self-induction, 806, 806–807
 mutual induction contrasted, 813
 semi-classical angle, 1116
 semiconducting devices, 696
 semiconductors, 563. *See also p-n junctions; transistors*
 doped, 1152–1153, 1153
 electrical properties, 1150–1151, 1151
 light-emitting diodes (LEDs), 1157–1158, 1158
 nanocrystallites, 1093, 1093
 resistivities of selected, 689t
 unit cell, 1142
 semimajor axis, of orbits, 342–343, 343
 planets of Solar System, 344t
 series circuits
 capacitors, 663–666, 664, 715t, 842–844, 843
 inductors, 842–844, 843
 RC, 720, 720–723

resistors, 710, 710, 715t, 842–844, 843
RLC, 833, 833–834, 842–846, 843
 summary of relations, 715t
 series limit, 1100, 1101
 shearing stress, 315, 315–316
 shear modulus, 316
 shells, 1106, 1115, 1115t
 and characteristic x-ray spectrum, 1128
 and energy levels in crystalline solids, 1143
 and periodic table, 1126–1127
 shell theorem, 332
 shock waves, 465, 465
 short wave radiation, 891
 side maxima
 diffraction patterns, 990, 990
 interference patterns, 966
 sievert (unit), 1184
 sigma particle, 1219
 and eightfold way, 1231t
 and strangeness, 1231
 sign
 acceleration, 20
 displacement, 13–14
 heat, 484
 velocity, 20, 27
 work, 143
 silicon
 effect of doping on Fermi energy, 1163
 electric properties compared to copper, 696t, 696–697, 1143t, 1151
 electron number density in conduction band, 1164
 germanium compared to, 1162
 in MOSFETs, 1159
 properties of *n*- and *p*-doped compared, 1153t
 resistivity of pure, 689t
 as semiconductor, 563, 696–697, 1150–1151
 temperature coefficient of resistivity, 1151
 unit cell, 1142, 1142
 silicon, *n*-type resistivity, 689t
 silicon, *p*-type resistivity, 689t
 silk
 rubbing glass rod with, 562, 562–562, 572
 simple harmonic motion, 386–389, 387, 388
 acceleration, 389, 389
 angular, 394, 394–395
 damped, 400, 400–401, 401
 energy, 392–394, 393

- force law for, 390–392
 pendulums, 395, 395–398,
 397
 and uniform circular motion, 398–399, 399
 velocity, 388, 388, 389, 389
 waves produced by, 413–414
 simple harmonic oscillators.
See also pendulums
 angular, 394, 394–395
 linear, 390, 390–392
 simple magnifying lens, 943,
 943
 simple pendulum, 395,
 395–396
 simultaneity
 consequences of Lorentz transformation equations, 1037
 relativity of, 1025–1026, 1026
 sine, 43, 43
 single-component forces, 88
 single-loop circuits, 836
 current calculation, 707,
 707–708
 internal resistance, 709, 709
 resistances in series, 710, 710
 single-slit diffraction, 963–964,
 964, 992, 992–995, 993
 intensity, 995–999, 996, 998
 sinusoidal waves, 414–417, 415,
 417
 SI units, 2
 skateboarding, motion analyzed, 67
 slab
 rotational inertia, 255t
 sliding block, 700, 100–101
 sliding friction, 118, 118
 slit diffraction, 990
 slope, of line, 15–16, 16
 Snell's law, 906, 959
 soap bubbles, interference patterns from, 973, 976, 976
 sodium, 1126, 1186
 sodium chloride, 1126
 index of refraction, 906t
 x-ray diffraction, 1012, 1012
 sodium doublet, 1140
 sodium vapor lamp, 1058
 soft reflection, of traveling waves at boundary, 432
 solar system, 1243
 solar wind
 and comets, 917, 917
 solenoids, 774
 induced emf, 794–795
 inductance, 805–806
 magnetic energy density, 812
 magnetic field, 776, 776–777
 solid bodies
 center of mass, 203–206
 solids
 compressibility, 318
 electrical properties of crystalline, 1142, 1142–1143
 heat absorption, 485–488
 specific heats of selected, 485t
 speed of sound in, 447t
 thermal conductivity of selected, 494t
 thermal expansion, 481–483,
 482
 solid state, 486
 solid-state electronic devices, 1142
 sonar, 445
 sonic boom, 465
 sound intensity, 452–456, 453
 sound level, 452–456
 selected sounds, 454t
 sound waves, 413–414
 beats, 459, 459–460
 Doppler effect, 461–464,
 462, 463
 interference, 451, 451–452
 sources of musical, 456–459,
 457, 458
 speed of, 446–448, 447t
 supersonic speed, 465, 465
 traveling waves, 448,
 448–450
 south pole, magnets, 738–739,
 739, 862, 862
 space charge, 1155
 space curvature, 347, 347
 space time, 347, 1057, 1241
 spacetime coordinates,
 1024–1025
 spark, *See* electric spark
 special theory of relativity, 87,
 893, 1023, 1029
 specific heat, 485. *See also* molar specific heat
 selected materials, 485t
 speckle, 969
 spectral emission lines, 1082
 spectrum, 1101
 speed. *See also* angular speed
 average in one-dimensional motion, 16
 one-dimensional motion, 18
 relating linear to angular, 250
 in rolling, 276–277, 277
 traveling waves, 417–420
 speed of light, 893
 as ultimate speed, 1023
 speed of light postulate, 1023–1024
 speed of sound, 446–448
 and RMS speed in gas, 513
 in various media, 447t
 speed parameter, in time dilation, 1028, 1028
 spherical aberration, 946
 spherical capacitors
 calculating capacitance, 661–662
 spherical conductors
 Coulomb's law for, 566,
 566–570
 spherically symmetric charge distribution, 635
 Gauss' law for, 619, 619
 spherical mirrors, 928,
 928–930, 929, 930
 focal points, 929–930, 930
 images from, 928, 928, 930,
 930–933, 946, 946–947
 spherical refracting surfaces, 934, 934–935
 images from, 934, 934–935,
 946, 946–947
 spherical shell
 Coulomb's law for, 566
 566–570
 rotational inertia, 255t
 spherical symmetry, Gauss' law, 619, 619–620
 spherical waves, 446
 spin, 1115, 1115t, 1219–1221,
 1220
 isospin, 1245
 nuclides, 1168t, 1172
 spin angular momentum,
 871–872, 1115, 1115t,
 1116–1117, 1117
 and spin quantum number,
 1115t
 spin-down electron state, 871,
 1117, 1120, 1120
 spin-flipping, 883, 1121, 1121
 spin magnetic dipole moment,
 871–872, 872,
 1116–1117, 1117
 diamagnetic materials,
 875–876
 ferromagnetic materials,
 875, 879
 paramagnetic materials, 875,
 877–878
 spin magnetic quantum number, 871, 1115, 1115t,
 1219–1220
 spin quantum number, 1115,
 1115t, 1219–1220
 spin-up electron state, 871,
 1117, 1120, 1120
 spontaneous emission, 1132,
 1132–1133
 spontaneous otoacoustic emission, 470
 spring constant, 149
 spring force, 149–150
 as conservative force, 168,
 168
 work done by, 149, 149–151
 spring scale, 96, 96
 sprites, 585, 585
 s subshells, 1125, 1126
 stable equilibrium potential energy curves, 178
 stable static equilibrium, 305–306, 306
 stainless steel
 thermal conductivity, 494t
 standard kilogram, 6, 6
 standard meter bar, 3
 Standard Model, of fundamental particles, 1220–1222
 standards, 1
 standing waves, 431, 431–432,
 1083
 reflections at boundary, 432,
 432
 and resonance, 433, 433–435
 stars, 1057. *See also* black holes; neutron stars
 orbiting, 355
 Doppler shift, 1040
 formation in early universe, 1241
 fusion in, 1172, 1207,
 1208–1210, 1209
 preponderance of matter over antimatter, 1222
 rotational speed as function of distance from galactic center, 1240, 1240
 S2 star at center of Milky Way galaxy, white dwarfs, 341t, 360t
 state, 486
 state function
 entropy as, 539–540
 state properties, 537–538
 static equilibrium, 305,
 305–306, 306
 examples, 309–312, 313–311
 fluids, 362, 362
 indeterminate structures, 314, 314
 requirements of, 306–307
 static frictional force, 116–118,
 117
 rolling wheel, 278
 statistical mechanics, 550–553
 steady flow, 371
 steady-state current, 682
 series *RLC* circuits, 842

Stefan-Boltzmann constant, 496
 step-down transformer, 851
 step-up transformer, 851
 Stern-Gerlach experiment, 1118, 1118–1120
 stick-and-slip, 118
 stimulated emission, 1132–1133
 from comets, 1140, 1140
 Stirling engines, 546, 546
 Stirling's approximation, 553
 stopping potential, photoelectric effect, 1059–1061, 1061
 straight line charge distributions, 590
 straight-line motion, *See one-dimensional motion*
 strain, 315–318, 516
 strain gage, 316, 316
 strangeness, 1231
 and hypercharge, 1245
 strange particles, 1230–1231, 1231
 strange quark, 1232, 1233t, 1233
 streamlines
 in electric fields, 685
 in fluid flow, 371, 372
 stress, 315, 315
 stress-strain curves, 316, 316
 stress-strain test specimen, 316
 stretched strings, 446
 energy and power of traveling wave on, 421, 421–423
 harmonics, 434–435
 resonance, 433, 433–435
 standing waves, 431, 431–432, 432
 transverse and longitudinal waves on, 414, 414
 wave equation, 423–424
 wave speed on, 420–421, 420
 strike-slip, 57
 strokes, 543
 strong force, 115, 1173, 1221
 conservation of strangeness, 1231
 messenger particle, 1235–1236
 strong interaction, 1224
 strong nuclear force, 572
 subcritical state, nuclear reactors, 1203
 submarines
 Project Seafarer, 915
 sonar, 445
 subshells, 1106, 1115, 1115t

and energy levels in crystalline solids, 1143
 and periodic table, 1125
f subshells, 1125
 substrate, 1159
 subtraction
 of vectors by components, 45–46
 of vectors geometrically, 39–40, 40
 Sun, 1243
 convection cells in, 496
 density at center of, 360t, 1217
 escape speed, 341t
 fusion in, 1172, 1207, 1209–1210, 1209
 monitoring charged particles from, 682
 neutrinos from, 1181
 period of revolution about galactic center, 354
 pressure at center of, 361t
 radiation from, 889
 randomly polarized light, 901
 speed distribution of photons in core, 517
 superconductivity, 697
 superconductors, 563, 697, 697–698
 supercooling, 557
 supercritical state, nuclear reactors, 1203
 superluminal jets, galaxies, 1054
 supermassive black holes, 330
 supernova, 82, 341t, 1210, 1210, 1242–1243
 supernova SN1987a, 1181, 1181, 1210, 1210
 supernumeraries, 962, 962
 superposition principle for forces, *See principle of superposition*
 surface charge density, 586, 586t
 S waves, 468
 symmetric lenses, 938–939
 symmetry
 cylindrical, Gauss' law, 615, 615–616
 importance in physics, 605
 line, plane, point: center of mass of bodies with, 203
 planar, Gauss' law, 617, 617–618
 spherical, Gauss' law, 619, 619–620
 system, 92, 483
 systems of particles, *See particle systems*

T
 tangent, 43, 43
 tangential component
 of linear acceleration, 250–251
 of torque, 258
 target
 collisions in two dimensions, 224, 224
 elastic collisions in one dimension with moving, 222–223, 223
 elastic collisions in one dimension with stationary, 221, 221
 inelastic collisions in one dimension, 218
 series of collisions, 213, 213
 single collision, 211–212
 tattoo inks, magnetic particles in, 861, 861
 tau neutrinos, 1227, 1227t
 tau particle, 1227, 1227t
 telescopes, 943
 refracting, 945, 945–946
 surveillance satellites, 1017
 television, 869
 magnet applications, 735
 television waves, 413
 temperature, 476
 defined, 478
 and heat, 483
 and ideal gas law, 508–511
 measuring, 478–479
 and RMS speed of ideal gas, 511–513
 scalar nature of, 38
 as state property, 537
 work done by ideal gas at constant, 509, 509–510
 and zeroth law of thermodynamics, 478, 477
 temperature coefficient of resistivity, 691, 1143
 selected materials, 689t
 semiconductors, 1151
 silicon and copper compared, 696t, 1143t
 temperature field, 580
 temperature scales, 476, 476, 479–481
 compared, 480
 temporal separation, of events, 1027
 tensile stress, 316, 316
 tension force, 97–98, 98
 and elasticity, 316–317
 and wave speed on stretched string, 421
 terminals, batteries, 658, 706
 terminal speed, 122, 122–123
 tesla (unit), 737
 test charge, 580–581, 581
 Tevatron, 1234
 thermal agitation
 ferromagnetic materials, 879
 paramagnetic materials, 877
 thermal capture, of neutrons in nuclear reactor, 1202
 thermal conduction, 494–495, 495
 thermal conductivity, 494–495
 selected materials, 494t
 thermal conductor, 494
 thermal efficiency
 Carnot engines, 545–546
 Stirling engines, 546
 thermal energy, 168, 475
 and conservation of total energy, 183
 thermal equilibrium, 477
 thermal expansion, 481–483, 482
 thermal insulator, 494
 thermal neutrons, 1196–1201, 1202
 thermal radiation, 496–497
 thermal reservoir, 488, 489
 thermal resistance to conduction, 494
 thermodynamic cycles, 489, 490, 492
 thermodynamic processes, 488–491, 489
 graphical summary, 528
 thermodynamics, 477. *See also entropy; irreversible processes*
 first law, 491–494
 second law, 541–542
 zeroth law, 477, 477
 thermometers, 477
 constant-volume gas, 478, 478–479
 liquid-in-glass, 482
 thermonuclear bomb, 1211
 thermonuclear fusion, 1045, 1172, 1207–1208
 controlled, 1211–1213
 in Sun and stars, 1207, 1209–1210, 1209
 thermopiles, 706
 thermoscope, 477, 477
 thin films
 interference, 973, 973–979
 thin-lens approximation, 948
 thin-lenses, 936–942

- images from, 938, 938–942, 939, 947, 947–948
two-lens systems, 939–941, 940
- thin-lens formulas, 936
Newtonian and Gaussian forms contrasted, 956
- third-law force pair, 98
- three-dimensional electron traps, 1095, 1095–1096
- three-dimensional motion
acceleration, 61, 61
position and displacement, 58, 58–60
velocity, 60, 60–62, 61
- three-slit diffraction gratings, 1019
- thrust, 226
- thunderstorm sprites, 585, 585
- time
directional nature of, 536
relativity of, 1027, 1027–1031, 1037
scalar nature of, 38
units, 5–6
- time constant
LC oscillations, 826
RC circuits, 721, 721
RL circuits, 809
- time dilation, 1028
consequences of Lorentz transformation equations, 1037
tests of, 1029
and transverse Doppler effect, 1041, 1041–1042
- time intervals, 5, 5t
between relativistic events, 1027, 1027–1031
- time signals, 6
- tokamak, 1212
- ton, 11
- top quark, 1233t, 1234
- toroids
magnetic field, 777, 777–778
- torque, 258–260, 289t
and angular momentum of system of particles, 288
and conservation of angular momentum, 290
- on current loop, 752, 752–753
and gyroscope precession, 294, 295
- internal and external, 288
net (resultant), 259
- Newton's second law in angular form, 285
particle about fixed point, 281, 281–282
rolling down ramp, 278–279
- torr, 361
- torsion constant, torsion pendulum, 394
- torsion pendulum, 394, 394
- total internal reflection, 911, 911–912
- tour jeté, 291, 291
- tracer, for following fluid flow, 371, 371
- trajectory, in projectile motion, 67
- transformers, 850–852
autotransformer, 859, 859
energy transmission requirements, 850
ideal, 850, 850–852
impedance matching, 852
in *LC* oscillators, 891, 891
- transient current series *RLC* circuits, 844
- transistors, 697, 1059
field-effect-transistor (FET), 1159, 1159–1160
- MOSFET (metal-oxide-semiconductor-field-effect-transistor), 1159, 1159
- transition elements
paramagnetism, 875
- translation, 241. *See also* one-dimensional motion; two-dimensional motion
in rolling, 275–277, 276
- translational equilibrium, 306
- translational kinetic energy
ideal gases, 513–514
of rolling, 277
yo-yo, 280–281
- translational variables, 289t
- transmission coefficient, barrier tunneling, 1075
- transparent material, 906
- transuranic nuclides, 1204
- transverse Doppler effect, 1041, 1041–1042
- transverse motion, 414
- transverse waves, 413–414, 414
electromagnetic waves, 891
- traveling electromagnetic waves, 890–896, 893, 894
- traveling waves, 414, 1083
energy and power, 421, 421–423
sound, 448, 448–450
speed, 417, 417–419
wave function, 1071–1072
- triangular prism, 908, 908
angle of minimum deviation, 920
- trig functions, 43, 43
- triple point cell, 478, 478
- triple point of water, 478–479
- tritium, 1211, 1212
and valley of the nuclides, 1182
- triton, 1211
- tube length, compound microscope, 944
- tunnel diode, 1074
- tunneling, barrier, 1074, 1074–1076
- turbulent flow, 371
- turning points, in potential energy curves, 176–178, 177
- turns
in coils, 752–753
in solenoids, 776
- turns ratio, transformers, 852
- two-dimensional collisions, 224, 224
- two-dimensional electron traps, 1095, 1095–1096
- two-dimensional explosions, 216, 216
- two-dimensional motion
acceleration, 62–64, 63
position and displacement, 58–60, 58, 59
relative, 74–75, 74
velocity, 60–62, 60, 61
- U**
- ultimate strength, 316, 316
selected engineering materials, 317t
- ultrarelativistic proton, 1047
- ultrasound
bat navigation using, 464
blood flow speed measurement using, 473
- ultrasound imaging, 445, 445
- ultraviolet light, 413
- ultraviolet radiation, 869
- uncertainty principle, 1072, 1073, 1236
- uniform charge distributions
electric field lines, 581, 581–582
types of, 590
- uniform circular motion, 70–72, 71
centripetal force in, 124–128, 125
and simple harmonic motion, 398–399, 399
- uniform displacement current, 885
- uniform displacement-current density, 885
- uniform electric field, 582, 582, 884
- uniform electric flux, 884
- unit cell, 1012, 1012
metals, insulators, and semiconductors, 1142–1143, 1142
- United States Naval Observatory time signals, 6
- units, 1–2
changing, 3
length, 3–4
mass, 6–8
time, 5–6
- unit vectors, 44, 44
- universe
Big Bang, 1237, 1240–1243, 1241
- color-coded image of universe at 379,000 yrs old, 1242, 1242
- cosmic background radiation, 1239, 1241, 1243, 1243
- critical density for continued expansion, 1246
- dark energy, 1243
- dark matter, 1240
- estimated age, 1238
- expansion, 1238, 1243, 1242–1243, 1246
- temperature of early, 477
- unoccupied levels, 1122, 1144, 1185
- unpolarized light, 901, 901
- unstable equilibrium, 178
potential energy curves, 178
- unstable static equilibrium, 306
- up quark, 578, 655, 1232, 1233t, 1233, 1234
- uranium
density of nucleus, 360t
electric field at surface of nucleus, 580t
mass energy, 1043t
- uranium₂₂₈
alpha decay, 1177
half-life, 1178, 1179t
- uranium₂₃₅, 1215
enriching fuel, 1202, 1216
fission, 1196–1201, 1199, 1214
- fissionability, 1200t, 1201, 1206
- in natural nuclear reactor, 1206
- percentage in uranium ore, 1216
- spontaneous fission, 1214
- uranium₂₃₆, 1197, 1200t

uranium₂₃₈, 572, 1174
abundance in granite, 1192
alpha decay, 1178, 1177–1179
binding energy per nucleon,
1172
fissionability, 1200t, 1201,
1206
half-life, 1179, 1179t, 1192
uranium₂₃₉, 1200t
uranium enrichment, 1202,
1216

V

vacant levels, 1144
valence band, 1151, 1151
valence electrons, 1083, 1126,
1145
valence number, 1152
valley of nuclides, 1182, 1182
vaporization, 486
vapor state, 486
variable capacitor, 678
variable force
work done by general variable, 151–153, 152
work done by spring force,
149, 149–151
variable-mass systems, rockets,
224–226, 225
variable resistors, 727
vector angles, 41, 41
degrees and radians, 43
measuring, 43
vector-capable calculator, 44
vector components, 41, 41–43,
42
addition, 44–46
vector equation, 39
vector fields, 580
vector product, 48, 50, 49–52
vector quantities, 14, 38
vectors, 580
adding by components,
44–46
adding geometrically, 39,
39–40, 40
coupled, 1114
cross product, 49–52, 51
defined, 38, 38–40, 39
dot product, 48, 48
and laws of physics, 47
multiplication, 47–52, 48
resolving, 41
scalar product, 48–49
unit, 44, 44
vector product, 47, 50, 50–51
vector sum (resultant), 39,
39–40
velocity, 263t. *See also* acceleration; angular velocity;
forces; kinetic energy

graphical integration in motion analysis, 27, 27–28
line of sight, 355
and Newton's first law, 87–91
and Newton's second law,
91–94
one-dimensional motion,
14–18
projectile motion, 64–70
relative motion in one dimension, 73–74
relative motion in two dimensions, 74–75
relativity of, 1039, 1039
rockets, 224–226
sign of, 20
simple harmonic motion,
388, 388, 389, 389
two- and three-dimensional motion, 60–62, 61, 62
uniform circular motion,
70–72, 71
as vector quantity, 38
velocity amplitude
forced oscillations, 402, 402
simple harmonic motion,
389
velocity field, 606, 606
velocity vectors, 38
venturi meter, 383
vertical motion, in projectile motion, 66–67, 67
virtual focal point, 929, 929
virtual images
defined, 924
spherical mirrors, 930
spherical refracting surfaces,
934, 934
thin lenses, 938, 938
virtual photons, 1235
viscous drag force, 371
visible light, 413, 890
ultimate speed, 1023
volt, 630
voltage. *See also* potential difference
ac circuits, 842t
transformers, 850–852
voltage law, Kirchoff's, 708
volt-ampere, 695
voltmeters, 720, 720
volume
and ideal gas law, 508–511
as state property, 537
work done by ideal gas at constant, 510
volume charge density, 576,
578, 586, 586t
volume expansion, 482–483
volume flow rate, 373, 605–606
volume flux, 605

volume probability density,
1103, 1103, 1105, 1105,
1106

W

water
boiling and freezing points
in Celsius and Fahrenheit, 480t
bulk modulus, 317, 446
density, 360t
dielectric properties, 669t
diffraction of waves in tank,
963, 963
as electric dipole, 594, 594,
638
electric field within, 580t
heats of transformation, 486,
487
heavy, 1217
index of refraction, 906t
in microwave cooking, 595
as moderator for nuclear reactors, 1202
as polar dielectric, 671
polarization of light by reflection, 912
pure, as insulator, 563
RMS speed at room temperature, 512t
specific heats, 485t
speed of sound in, 447, 447t
tap water, as conductor, 563
thermal properties, 482–483
triple point, 478
water waves, 413
watt (W), 2, 155–156
Watt, James, 155
wave equation, 423–424
wave forms, 414, 414
wavefronts, 445, 445
wave function, 1071–1072. *See also* Schrödinger's equation
electron traps, 1088–1091,
1089
hydrogen ground state,
1102–1103, 1103
normalizing, 1089
wave interference, 425,
425–428, 426
sound waves, 451, 451–452
wavelength, 415, 415
and index of refraction,
960–962
sound waves, 449
wave on stretched string, 420
wavelength Doppler shift, 1040
waves. *See also* seismic waves;
sound waves; standing waves; stretched strings;
traveling waves
phasors, 428–430, 429
principle of superposition
for, 425, 425
transverse and longitudinal,
413–414, 414
types of, 413
wavelength and frequency,
414–417
wave shape, 415
wave speed, 417, 417–420
sound waves, 449
on stretched string, 420,
420–421
wave trains, 1131
weak force, 1221
messenger particle, 1235
weak interaction, 1225
weber (unit), 793
weight, 95–96
apparent weight in fluids,
369–370
weightlessness, 124
well depth, finite well electron traps, 1091
whiplash injury, 28
white dwarfs, 341t
density of core, 360t
white light
chromatic dispersion, 908,
908
Wien's law, 1247
Wilkinson Microwave Anisotropy Probe (WMAP), 1242
windings, solenoid, 776
window glass
thermal conductivity, 494t
Wintergreen LifeSaver, blue flashes, 564
WMAP (Wilkinson Microwave Anisotropy Probe), 1242
W messenger particle, 1236
work, 263t
Carnot engines, 545
and conservation of mechanical energy,
173–175
and conservation of total energy, 183–186, 184
defined, 142–143
done by electrostatic force,
630–631
done by external force with friction, 180–182
done by external force without friction, 180
done by gravitational force,
146, 146–148

done by ideal gas at constant pressure, 510
done by ideal gas at constant temperature, 509, 509–510
done by ideal gas at constant volume, 510
done by spring force, 149, 149–151
done by variable force, 151–155, 152
done in lifting and lowering objects, 147, 147–148
done on system by external force, 180–182, 181
first law of thermodynamics, 491–493
and heat, 484, 488–490
and kinetic energy, 143, 142–146
net, 144

path-dependent quantity, 490
path independence of conservative forces, 168–170, 169
and potential energy, 167, 167–170
and power, 155–156, 156
and rotational kinetic energy, 262–264
signs for, 143
work, energy, and emf, 706–707
work function, 1060
working substance, 543
work-kinetic energy theorem, 144–146, 263t
variable force, 153–155

X

x component, of vectors, 41–42, 41, 42

xenon decay chain, 1197
xi-minus particle, 1230, 1231t, 1235
x-ray diffraction, 1011–1013, 1012
x rays, 413, 889, 890
characteristic x-ray spectrum, 1128–1129, 1129
continuous x-ray spectrum, 1127, 1127–1128
and ordering of elements, 1127–1130
radiation dosage, 1184

Y

y component, of vectors, 41–42, 41
yield strength, 316, 316
selected engineering materials, 317t

Young's double-slit interference experiment, 965, 965–967
single-photon, wide-angle version, 1067, 1067
single-photon version, 1065, 1066
Young's modulus, 316
selected engineering materials, 317t
yo-yo, 280–281, 281

Z

zero angular position, 242
zero-point energy, 1090
zeroth law of thermodynamics, 477, 477
zeroth-order line, 1006
Z messenger particle, 1236

Some Physical Constants*

Speed of light	c	$2.998 \times 10^8 \text{ m/s}$
Gravitational constant	G	$6.673 \times 10^{-11} \text{ N} \cdot \text{m}^2/\text{kg}^2$
Avogadro constant	N_A	$6.022 \times 10^{23} \text{ mol}^{-1}$
Universal gas constant	R	$8.314 \text{ J/mol} \cdot \text{K}$
Mass–energy relation	c^2	$8.988 \times 10^{16} \text{ J/kg}$ 931.49 MeV/u
Permittivity constant	ϵ_0	$8.854 \times 10^{-12} \text{ F/m}$
Permeability constant	μ_0	$1.257 \times 10^{-6} \text{ H/m}$
Planck constant	h	$6.626 \times 10^{-34} \text{ J} \cdot \text{s}$ $4.136 \times 10^{-15} \text{ eV} \cdot \text{s}$
Boltzmann constant	k	$1.381 \times 10^{-23} \text{ J/K}$ $8.617 \times 10^{-5} \text{ eV/K}$
Elementary charge	e	$1.602 \times 10^{-19} \text{ C}$
Electron mass	m_e	$9.109 \times 10^{-31} \text{ kg}$
Proton mass	m_p	$1.673 \times 10^{-27} \text{ kg}$
Neutron mass	m_n	$1.675 \times 10^{-27} \text{ kg}$
Deuteron mass	m_d	$3.344 \times 10^{-27} \text{ kg}$
Bohr radius	a	$5.292 \times 10^{-11} \text{ m}$
Bohr magneton	μ_B	$9.274 \times 10^{-24} \text{ J/T}$ $5.788 \times 10^{-5} \text{ eV/T}$
Rydberg constant	R	$1.097\,373 \times 10^7 \text{ m}^{-1}$

*For a more complete list, showing also the best experimental values, see Appendix B.

The Greek Alphabet

Alpha	A	α	Iota	I	ι	Rho	P	ρ
Beta	B	β	Kappa	K	κ	Sigma	Σ	σ
Gamma	Γ	γ	Lambda	Λ	λ	Tau	T	τ
Delta	Δ	δ	Mu	M	μ	Upsilon	Y	ν
Epsilon	E	ϵ	Nu	N	ν	Phi	Φ	ϕ, φ
Zeta	Z	ζ	Xi	Ξ	ξ	Chi	X	χ
Eta	H	η	Omicron	O	\o	Psi	Ψ	ψ
Theta	Θ	θ	Pi	Π	π	Omega	Ω	ω

Some Conversion Factors*

Mass and Density

$$\begin{aligned}1 \text{ kg} &= 1000 \text{ g} = 6.02 \times 10^{26} \text{ u} \\1 \text{ slug} &= 14.59 \text{ kg} \\1 \text{ u} &= 1.661 \times 10^{-27} \text{ kg} \\1 \text{ kg/m}^3 &= 10^{-3} \text{ g/cm}^3\end{aligned}$$

Length and Volume

$$\begin{aligned}1 \text{ m} &= 100 \text{ cm} = 39.4 \text{ in.} = 3.28 \text{ ft} \\1 \text{ mi} &= 1.61 \text{ km} = 5280 \text{ ft} \\1 \text{ in.} &= 2.54 \text{ cm} \\1 \text{ nm} &= 10^{-9} \text{ m} = 10 \text{ Å} \\1 \text{ pm} &= 10^{-12} \text{ m} = 1000 \text{ fm} \\1 \text{ light-year} &= 9.461 \times 10^{15} \text{ m} \\1 \text{ m}^3 &= 1000 \text{ L} = 35.3 \text{ ft}^3 = 264 \text{ gal}\end{aligned}$$

Time

$$\begin{aligned}1 \text{ d} &= 86400 \text{ s} \\1 \text{ y} &= 365 \frac{1}{4} \text{ d} = 3.16 \times 10^7 \text{ s}\end{aligned}$$

Angular Measure

$$\begin{aligned}1 \text{ rad} &= 57.3^\circ = 0.159 \text{ rev} \\\pi \text{ rad} &= 180^\circ = \frac{1}{2} \text{ rev}\end{aligned}$$

Speed

$$\begin{aligned}1 \text{ m/s} &= 3.28 \text{ ft/s} = 2.24 \text{ mi/h} \\1 \text{ km/h} &= 0.621 \text{ mi/h} = 0.278 \text{ m/s}\end{aligned}$$

Force and Pressure

$$\begin{aligned}1 \text{ N} &= 10^5 \text{ dyne} = 0.225 \text{ lb} \\1 \text{ lb} &= 4.45 \text{ N} \\1 \text{ ton} &= 2000 \text{ lb} \\1 \text{ Pa} &= 1 \text{ N/m}^2 = 10 \text{ dyne/cm}^2 \\&\quad = 1.45 \times 10^{-4} \text{ lb/in.}^2 \\1 \text{ atm} &= 1.01 \times 10^5 \text{ Pa} = 14.7 \text{ lb/in.}^2 \\&\quad = 76.0 \text{ cm Hg}\end{aligned}$$

Energy and Power

$$\begin{aligned}1 \text{ J} &= 10^7 \text{ erg} = 0.2389 \text{ cal} = 0.738 \text{ ft} \cdot \text{lb} \\1 \text{ kW} \cdot \text{h} &= 3.6 \times 10^6 \text{ J} \\1 \text{ cal} &= 4.1868 \text{ J} \\1 \text{ eV} &= 1.602 \times 10^{-19} \text{ J} \\1 \text{ horsepower} &= 746 \text{ W} = 550 \text{ ft} \cdot \text{lb/s}\end{aligned}$$

Magnetism

$$1 \text{ T} = 1 \text{ Wb/m}^2 = 10^4 \text{ gauss}$$

*See Appendix D for a more complete list.