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DETERMINATION METHOD, QUALITY ASSURANCE METHOD, ELECTROLYSIS DEVICE, AND ELECTROLYSIS METHOD

Abstract

A determination method determines whether or not target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules which include: hydrogen molecules produced by water electrolysis; or molecules produced using the hydrogen molecules as a raw material. In the determination method, the method includes determining that the target molecules are the electrolytic hydrogen-containing molecules when an abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than an abundance ratio of deuterium to light hydrogen in nature.

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Background/Summary

CROSS-REFERENCE TO RELATED APPLICATIONS [0001] This application is a continuation application of International Application No. PCT/JP2023/038103, filed on Oct. 20, 2023, which claims priority to Japanese Patent Application No. 2022-175572, filed on Nov. 1, 2022, the entire contents of which are incorporated by reference herein.

BACKGROUND

1. Technical Field

[0002] The disclosure relates to a determination method, a quality assurance method, an electrolysis device, and an electrolysis method.

2. Description of the Related Art

[0003] Carbon dioxide has been regarded as a cause of global warming, and there has been a worldwide movement to curb carbon dioxide emissions. Hydrogen is attracting attention as an alternative to fossil fuels because it does not emit carbon dioxide when used, and also because it can be obtained by electrolyzing water using renewable energy. As a method for electrolyzing water to produce hydrogen, an alkaline type water electrolysis device disclosed in WO 2019/181662 is known.

SUMMARY

[0004] Hydrogen is conventionally produced industrially by steam-reforming fossil fuels, such as natural gas. However, there is no way to confirm that any give sample of hydrogen has been produced by water electrolysis. Ammonia is expected to be used as a next-generation fuel, and hydrocarbons are used as raw materials for various chemical products. Molecules thereof can be produced using hydrogen as a raw material, but as with hydrogen, there is no method to confirm that hydrocarbons are produced by water electrolysis. When it is possible to confirm whether these molecules are produced via water electrolysis, the quality of these molecules can be assured.

[0005] Accordingly, it is an object of the present disclosure to provide a determination method and a quality assurance method that can confirm whether target molecules are: hydrogen which has been produced by water electrolysis; or molecules which have been produced using the hydrogen as a raw material. Also, it is an object of the present disclosure to provide an electrolysis device and an electrolysis method that can easily implement these methods.

[0006] A determination method according to the present disclosure is a determination method for determining whether or not target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules which include: hydrogen molecules produced by water electrolysis; or molecules produced using the hydrogen molecules as a raw material. The method includes determining that the target molecules are the electrolytic hydrogen-containing molecules when an abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than an abundance ratio of deuterium to light hydrogen in nature.

[0007] The target molecules may be hydrogen molecules, ammonia, or a hydrocarbon.

[0008] A quality assurance method according to the present disclosure is a quality assurance

method for assuring that target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules which include: hydrogen molecules produced by water electrolysis; or molecules produced using the hydrogen molecules as a raw material. The quality assurance method includes assuring that the target molecules are the electrolytic hydrogen-containing molecules when an abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than an abundance ratio of deuterium to light hydrogen in nature.

[0009] An electrolysis device according to the present disclosure includes: an electrolytic cell that electrolyzes water; a circulation flow path where water to be electrolyzed in the electrolytic cell circulates; a water supply flow path that supplies pure water to the circulation flow path; and a drainage flow path that drains part or all of the water in the circulation flow path, downstream of the electrolytic cell and upstream of the water supply via the water supply flow path. In the electrolysis device, an abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell is smaller than an abundance ratio of deuterium to light hydrogen in nature.

[0010] A flow rate control device that controls an amount of water drained in the circulation flow path may be provided in the drainage flow path.

[0011] Water supplied to the electrolytic cell may be alkaline water, and the electrolysis device may further include a membrane separator that is provided in the drainage flow path and may include a permeable membrane which selectively passes water in the alkaline water therethrough.

[0012] An electrolysis method according to the present disclosure includes: an electrolysis step of electrolyzing water in an electrolytic cell; a water supply step of supplying pure water to a circulation flow path that circulates water to be electrolyzed in the electrolytic cell; and a drainage step of draining part or all of the water in the circulation flow path, downstream of the electrolytic cell and upstream of the water supply in the water supply step. In the electrolysis method, an abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell is smaller than an abundance ratio of deuterium to light hydrogen in nature.

[0013] At least one selected from the group consisting of: a ratio of an amount of water consumed by electrolysis in the electrolytic cell to an amount of water supplied to the electrolytic cell; a ratio of an abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell to an abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the electrolytic cell; and a ratio of a flow rate of water drained in the drainage step to a flow rate of water discharged from the electrolytic cell, may be controlled.

[0014] A determination device includes a determination unit that determines whether or not target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules which include hydrogen molecules produced by water electrolysis, or molecules produced using the hydrogen molecules as a raw material. The determination unit determines that the target molecules are the electrolytic hydrogen-containing molecules when an abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than an abundance ratio of deuterium to light hydrogen in nature.

[0015] The determination device includes a measurement unit that measures the abundance ratio of deuterium to light hydrogen in the target molecules. The determination device includes a determination unit that determines that the target molecules are the electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules obtained by the measurement unit is less than or equal to a predetermined threshold which is less than the abundance ratio of deuterium to light hydrogen in nature. The determination device includes an output unit that outputs a determination result determined by the determination unit.

[0016] According to the present disclosure, it is possible to provide a determination method and a

quality assurance method that can confirm whether target molecules are hydrogen which has been produced by water electrolysis, or molecules produced using the hydrogen as a raw material. According to the present disclosure, it is also possible to provide an electrolysis device and an electrolysis method that can easily implement these methods.

Description

BRIEF DESCRIPTION OF THE DRAWINGS

[0017] FIG. 1 is schematic diagram illustrating an example of a low temperature type water electrolysis device according to the present embodiment.

[0018] FIG. 2 is a schematic diagram illustrating an example of a PEM type water electrolysis device.

[0019] FIG. 3 is a schematic diagram illustrating an example of an alkaline type water electrolysis device.

[0020] FIG. 4 is a schematic diagram illustrating an example of an AEM type water electrolysis device.

[0021] FIG. 5 is a schematic diagram illustrating an example of a medium-high temperature steam electrolysis device according to the present embodiment.

[0022] FIG. 6 is a schematic diagram illustrating an example of an SOEC type water electrolysis device.

[0023] FIG. 7 is a schematic diagram illustrating an example of a PCEC type water electrolysis device.

[0024] FIG. 8 is a schematic diagram illustrating an example of an electrolysis device including a permeation device.

DESCRIPTION OF THE EMBODIMENTS

[0025] Some exemplary embodiments will be described with reference to the drawings. Note that dimensional ratios in the drawings are exaggerated for convenience of the description and are sometimes different from actual ratios.

[0026] A determination method according to the present embodiment determines whether or not target molecules are electrolytic hydrogen-containing molecules. The electrolytic hydrogen-containing molecules include hydrogen molecules produced by water electrolysis, and molecules produced by using the hydrogen molecules as a raw material. Water electrolysis can be performed using renewable energy. When target molecules are electrolytic hydrogen-containing molecules, and water electrolysis is performed using renewable energy, it can be determined whether the target molecules are derived from renewable energy.

[0027] Target molecules are molecules including elemental hydrogen. The target molecules may be hydrogen molecules, ammonia, or a hydrocarbon. Similarly, electrolytic hydrogen-containing molecules may be hydrogen molecules, ammonia, or a hydrocarbon. Hydrogen and ammonia can be used as carbon-free fuels. Thus, instead of fossil fuels, hydrogen and ammonia can be used as fuels derived from renewable energy. Hydrocarbons can be produced using carbon dioxide as a raw material. Thus, carbon dioxide included in plant emission gas can be recovered, and the recovered carbon dioxide can be effectively used as a raw material for chemical products.

[0028] Hydrogen molecules can be produced by water electrolysis. Water electrolysis can be performed using an electrolysis device described below. The hydrogen molecules may be hydrogen gas. Molecules such as ammonia and a hydrocarbon can be produced using hydrogen molecules produced by water electrolysis as a raw material. Ammonia can be produced using hydrogen molecules as a raw material, for example, with the Haber-Bosch method. Hydrocarbons may include at least one of methane or olefin. Methane can be produced by a methanation reaction using hydrogen molecules as a raw material. Olefin (alkene) can be produced by a Fisher-Tropsch

reaction using hydrogen molecules as a raw material.

[0029] It is known that there are three isotopes of elemental hydrogen present in nature: light hydrogen (^1H or H), deuterium (^2H or D), and tritium (tritium: ^3H or T). Light hydrogen is the most abundant isotope of elemental hydrogen in nature. Deuterium is a stable isotope of elemental hydrogen. Tritium is a radioactive isotope, and the amount of tritium present in nature is very small.

[0030] In the determination method according to the present embodiment, target molecules are determined to be electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than the abundance ratio of deuterium to light hydrogen in nature. Electrolytic hydrogen-containing molecules produced using a method described below have a smaller abundance ratio of deuterium. Therefore, when the abundance ratio of deuterium in target molecules is smaller than that in nature, the target molecules can be determined to be electrolytic hydrogen-containing molecules.

[0031] Specifically, in the determination method according to the present embodiment, deuterium, which is a stable isotope of elemental hydrogen, is used as a tracer to determine whether or not target molecules are electrolytic hydrogen-containing molecules. Although details will be described below, deuterium molecules, such as HD and D_2 , have a slower reaction rate than light hydrogen molecules, such as H_2 . Thus, the abundance ratio of deuterium to light hydrogen in hydrogen molecules obtained by water electrolysis using this reaction rate difference is smaller than the abundance ratio of HDO and D_2O to H_2O in electrolytic cell supply water. Therefore, it is possible to determine that target molecules are electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0032] The abundance ratio of deuterium to light hydrogen in hydrogen molecules obtained by water electrolysis as described above is smaller than that in nature. Thus, hydrogen-containing molecules, such as ammonia or a hydrocarbon, made using hydrogen molecules produced by water electrolysis as a raw material have a similar abundance ratio of deuterium as the hydrogen molecules. In contrast, the abundance ratio of deuterium to light hydrogen in hydrogen obtained by steam-reforming fossil fuels, such as natural gas, is similar to the abundance ratio of deuterium to light hydrogen in nature. Therefore, the abundance ratio of deuterium in hydrogen-containing molecules, such as ammonia or a hydrocarbon produced by using hydrogen molecules as a raw material having a smaller deuterium ratio than that in nature, is also smaller than the ratio of deuterium in nature.

[0033] The abundance ratio of deuterium to light hydrogen in target molecules can be obtained by calculating a molar ratio of deuterium to light hydrogen included in the target molecules. Specifically, the abundance ratio of deuterium to light hydrogen in target molecules is a molar ratio of molecules including at least one deuterium atom to molecules including only light hydrogen atoms among molecules included in the target molecules. The abundance ratio of deuterium can be obtained using a mass spectrometer. The abundance ratio of deuterium can also be obtained using a mass spectrometer combined with a separation device, such as a gas chromatograph. The abundance ratio of deuterium can also be obtained using a gas chromatograph combined with a detector, such as a TCD (thermal conductivity detector).

[0034] The abundance ratio of deuterium to light hydrogen in nature is said to be 184 ppm or less. Thus, the abundance ratio of deuterium to light hydrogen in nature may be, for example, 184 ppm or less. The abundance ratio of deuterium to light hydrogen in nature may be the abundance ratio of deuterium to light hydrogen in Vienna Standard Mean Ocean Water (VSMOW). The abundance ratio of deuterium to light hydrogen in Vienna Standard Mean Ocean Water is about 155 ppm.

[0035] The above threshold may be smaller than the abundance ratio of deuterium to light

hydrogen in nature. The threshold may be, for example, 120 ppm, 100 ppm, 80 ppm, 60 ppm, 40 ppm, 20 ppm, or 10 ppm. When the abundance ratio of deuterium is small, it can be easily determined whether target molecules are electrolytic hydrogen-containing molecules. Note that the threshold may be greater than 0 ppm.

[0036] As described above, the determination method according to the present embodiment determines whether target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules including: hydrogen molecules produced by water electrolysis; or molecules produced using the hydrogen molecules as a raw material. The determination method determines that target molecules are electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0037] As described above, the abundance ratio of deuterium is smaller for: hydrogen molecules obtained by water electrolysis using a reaction rate difference; or molecules produced using the hydrogen molecules as a raw material. Thus, it can be determined that target molecules are electrolytic hydrogen-containing molecules when the abundance ratio of deuterium in the target molecules is smaller than that in nature. Thus, in the determination method according to the present embodiment, it is possible to confirm whether target molecules are hydrogen produced by water electrolysis, or molecules produced using the hydrogen as a raw material.

[0038] Water electrolysis can be performed using renewable energy, and it is possible to determine that target molecules are molecules produced using renewable energy in the determination method according to the present embodiment. That is, it becomes possible to construct traceability of molecules produced using renewable energy, by measuring the abundance ratio of deuterium to light hydrogen in target molecules. The method according to the present embodiment is particularly useful for sampling inspection at the time of receiving goods. Since the method according to the present embodiment is useful for sampling inspection, the quality can be assured by attaching an analysis result of produced electrolytic hydrogen-containing molecules, to a product as a quality record.

[0039] That is, the method according to the present embodiment may be a quality assurance method which assures that target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules including hydrogen molecules produced by water electrolysis, or molecules produced using the hydrogen molecules as a raw material. The quality assurance method may assure that target molecules are electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0040] In the quality assurance method according to the present embodiment, the quality of target molecules can be confirmed by analyzing the abundance ratio of deuterium to light hydrogen in the target molecules when the target molecules are received. In the quality assurance method according to the present embodiment, the quality of target molecules to be shipped can be assured by analyzing the abundance ratio of deuterium to light hydrogen in the target molecules before shipping the target molecules. The quality of target molecules may be attached to products as an assurance certificate or a label

[0041] Note that the determination method may perform determination using a determination device including a determination unit. The determination device may include, for example, a measurement unit, a determination unit, and an output unit. The measurement unit may include a device for measuring the abundance ratio of deuterium to light hydrogen in target molecules. The measurement unit may include a mass spectrometer, for example. The measurement unit may be a combination of a mass spectrometer and a separation device, such as a gas chromatograph. The measurement unit may include a combination of a gas chromatograph and a detector.

[0042] The determination unit determines whether target molecules including elemental hydrogen

are electrolytic hydrogen-containing molecules including: hydrogen molecules produced by water electrolysis; or molecules produced using the hydrogen molecules as a raw material. The determination unit determines that target molecules are electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than the abundance ratio of deuterium to light hydrogen in nature. The determination unit may determine that target molecules are electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules obtained by the measurement unit is less than or equal to a predetermined threshold which is smaller than the abundance ratio of deuterium to light hydrogen in nature. An abundance ratio of deuterium data signal which has been output from the measurement unit is output to the determination unit, and the determination unit may acquire the data output from the measurement unit. The determination unit may be a CPU (central processing unit), or a computer including a memory, for example. The CPU reads a determination program stored in the memory, and can determine whether target molecules are electrolytic hydrogen-containing molecules, based on the abundance ratio of deuterium to light hydrogen in the target molecules, obtained by the measurement unit, and a threshold. The output unit outputs a determination result determined by the determination unit. Examples of the output unit include a monitor and a printer. For example, the output unit can output a determination result that target molecules are electrolytic hydrogen-containing molecules, or target molecules are not electrolytic hydrogen-containing molecules, to the output unit.

(Electrolysis Device)

[0043] Next, an electrolysis device according to the present embodiment will be described. The electrolysis device according to the present embodiment can perform the water electrolysis described in the above embodiment. The electrolysis device according to the present embodiment may be a low temperature type water electrolysis device, or a medium-high temperature steam electrolysis device.

(Low Temperature Type Water Electrolysis Device)

[0044] First, an example of a low temperature type water electrolysis device will be described with reference to FIG. 1. As illustrated in FIG. 1, an electrolysis device 1 according to the present embodiment includes an electrolytic cell 10, a circulation flow path 20, a water supply flow path 30, and a drainage flow path 40.

[0045] The electrolytic cell 10 electrolyzes water. Hydrogen and oxygen are produced by electrolysis of water. An electrolysis process in the electrolytic cell 10 may be alkaline type water electrolysis, solid polymer type water electrolysis, or a combination thereof. The electrolysis process in the electrolytic cell 10 may be PEM (proton exchange membrane) type water electrolysis, alkaline type water electrolysis, AEM (anion exchange membrane) type water electrolysis, or the like.

[0046] The electrolytic cell 10 includes a membrane 11, a cathode 12, and an anode 13. The electrolytic cell 10 includes a DC power supply (not illustrated) electrically connected to the cathode 12 and the anode 13, and water is electrolyzed by applying a voltage to the cathode 12 and the anode 13.

[0047] Water electrolyzed in the electrolytic cell 10 circulates in the circulation flow path 20. Since pure water is usually used as water supplied to the electrolysis device 1, pure water can be effectively utilized by circulating the water. The water supply flow path 30 and the drainage flow path 40 are connected to the circulation flow path 20. The water supply flow path 30 supplies pure water to the circulation flow path 20. The pure water may be water having an electrical resistivity of 0.1 $\Omega \cdot \text{cm}$ or more at 25° C. The electrical resistivity of pure water may be 20 $\text{M}\Omega \cdot \text{cm}$ or less, 10 $\text{M}\Omega \cdot \text{cm}$ or less, or 1.5 $\text{M}\Omega \cdot \text{cm}$ or less. The drainage flow path 40 drains part or all of the water in the circulation flow path 20 downstream of the electrolytic cell 10, and upstream of water supply via the water supply flow path 30. The drainage flow path 40

may be provided with a flow rate control device **41** for controlling the amount of water drained in the circulation flow path **20**. The flow rate control device **41** can control the flow rate of water flowing in the drainage flow path **40**, and the amount of water in the circulation flow path **20** drained from the drainage flow path **40** can be controlled by the flow rate control device **41**. Thus, the abundance ratio of deuterium in hydrogen molecules produced in the electrolytic cell **10** can be adjusted. The flow rate control device **41** may be a flow rate control valve or the like.

[0048] The circulation flow path **20** may include a cathode-side water supply pipe **21**, an anode-side water supply pipe **22**, a cathode-side drain pipe **23**, and an anode-side drain pipe **24**. The electrolysis device **1** may include an electrolytic solution supply tank **50** provided in the circulation flow path **20**, a hydrogen gas-liquid separator **60** provided in the circulation flow path **20**, and an oxygen gas-liquid separator **65** provided in the circulation flow path **20**. The cathode-side water supply pipe **21** is provided with a pump **25**. The anode-side water supply pipe **22** is provided with a pump **26**. The cathode-side drain pipe **23** is provided with the hydrogen gas-liquid separator **60**. The anode-side drain pipe **24** is provided with the oxygen gas-liquid separator **65**.

[0049] Supplemental water is supplied to the electrolytic solution supply tank **50** via the water supply flow path **30**, and water to be electrolyzed in the electrolytic cell **10** is stored in the electrolytic solution supply tank **50**. An outlet of the electrolytic solution supply tank **50** is connected to an inlet of the electrolytic cell **10** on a cathode **12** side via the cathode-side water supply pipe **21**. Water is supplied from the electrolytic solution supply tank **50** to the cathode **12** side of the electrolytic cell **10** by operating the pump **25**. The outlet of the electrolytic solution supply tank **50** is connected to an inlet of the electrolytic cell **10** on an anode **13** side via the anode-side water supply pipe **22**. Water is supplied from the electrolytic solution supply tank **50** to the anode **13** side of the electrolytic cell **10** by operating the pump **26**.

[0050] The outlet of the electrolytic cell **10** on the cathode **12** side is connected to the inlet of the electrolytic solution supply tank **50** via the cathode-side drain pipe **23**. The cathode-side drain pipe **23** is provided with the hydrogen gas-liquid separator **60**. The outlet of the electrolytic cell **10** on the anode **13** side is connected to the inlet of the electrolytic solution supply tank **50** via the anode-side drain pipe **24**. The anode-side drain pipe **24** is provided with the oxygen gas-liquid separator **65**. Water which has passed through the electrolytic cell **10** is supplied to the hydrogen gas-liquid separator **60** together with hydrogen gas produced at the cathode **12**, and to the oxygen gas-liquid separator **65** together with oxygen gas produced at the anode **13**.

[0051] In the hydrogen gas-liquid separator **60**, hydrogen produced by electrolysis at the cathode **12**, and water discharged without being electrolyzed in the electrolytic cell **10** are separated. The hydrogen separated by the hydrogen gas-liquid separator **60** is recovered and stored, for example, in a storage tank. In contrast, the water separated by the hydrogen gas-liquid separator **60** is supplied to the electrolytic solution supply tank **50** via the cathode-side drain pipe **23**.

[0052] In the oxygen gas-liquid separator **65**, oxygen produced by electrolysis at the anode **13**, and water discharged without being electrolyzed in the electrolytic cell **10** are separated. The oxygen separated by the oxygen gas-liquid separator **65** is stored, for example, in a storage tank. In contrast, the water separated by the oxygen gas-liquid separator **65** is supplied to the electrolytic solution supply tank **50** via the anode-side drain pipe **24**.

[0053] Water discharged from the electrolytic cell **10** without being electrolyzed in the electrolytic cell **10** is stored in the electrolytic solution supply tank **50**, and the water circulates between the electrolytic cell **10** and the electrolytic solution supply tank **50**.

[0054] The electrolysis device **1** may include a control unit **70**. The control unit **70** may be electrically connected to at least one selected from the group consisting of the electrolytic cell **10**, the pump **25**, the pump **26**, and the flow rate control device **41**. The control unit **70** may control at least one of the applied voltage or the current density of the electrolytic cell **10**. The control unit **70** may control the flow rate of water supplied to the electrolytic cell **10** by operating at least one of the pump **25** or the pump **26**. The control unit **70** may control the flow rate of water in the

circulation flow path **20** drained from the drainage flow path **40** by operating the flow rate control device **41**. With these controls, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the electrolytic cell **10** can be adjusted.

[0055] Next, electrolytic reactions in a PEM type water electrolysis device, an alkaline type water electrolysis device, and an AEM type water electrolysis device will be described in detail.

(PEM Type Water Electrolysis Device)

[0056] First, an example of a PEM type water electrolysis device will be described with reference to FIG. 2. As illustrated in FIG. 2, in the PEM type water electrolysis device, water is supplied to the anode **13** via the anode-side water supply pipe **22** of the electrolytic cell **10**. At the anode **13**, oxygen and hydrogen ions (H^+) are produced from water by electrolysis. The membrane **11** is a PEM, and hydrogen ions (H^+) pass through the membrane **11** to move from the anode **13** side to the cathode **12** side. At the cathode **12**, hydrogen gas is produced from hydrogen ions that have passed through the membrane **11**. Water may be supplied to the cathode **12** side via the cathode-side water supply pipe **21**. Water need not be supplied via the cathode-side water supply pipe **21**.

[0057] The rate at which deuterium ions (D^+) pass through the membrane **11** is slower than the rate at which light hydrogen ions (H^+) pass through the membrane **11**. At the anode **13**, the rate at which D^+ is produced from HDO and D_2O is slower than the rate at which H^+ is produced from H_2O . Thus, at the cathode **12**, the amount of deuterium gas produced, such as HD gas and D_2 gas, is smaller than the amount of light hydrogen gas produced. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the PEM type water electrolysis device is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell **10**.

(Alkaline Type Water Electrolysis Device)

[0058] Next, an example of an alkaline type water electrolysis device will be described with reference to FIG. 3. As illustrated in FIG. 3, in the alkaline type water electrolysis device, water is supplied to the cathode **12** and the anode **13** of the electrolytic cell **10** via the cathode-side water supply pipe **21** and the anode-side water supply pipe **22**, respectively. At the cathode **12**, hydrogen and hydroxide ions (OH^-) are produced from water by electrolysis. Hydroxide ions (OH^-) pass through the membrane **11** to move from the cathode **12** side to the anode **13** side. At the anode **13**, oxygen is produced from hydroxide ions (OH^-) which have passed through the membrane **11**. The membrane **11** is a diaphragm which may include at least one selected from the group consisting of polysulfone, PTFE (polytetrafluoroethylene), asbestos, polyolefin, and an anion exchange membrane (AEM). The anion exchange membrane may be a resin having a quaternary ammonium group and an imidazolium group. Alkaline water passing through the electrolytic cell **10** may include an aqueous solution of an alkali metal hydroxide. The alkali metal hydroxide may include at least one of sodium hydroxide or potassium hydroxide.

[0059] The rate at which OD^- , HD gas, and D_2 gas are produced from HDO and D_2O at the cathode **12** is slower than the rate at which OH^- and H_2 gas are produced from H_2O . Thus, at the cathode **12**, the amount of deuterium gas produced, such as HD gas and D_2 gas, is smaller than the amount of light hydrogen gas generated. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the alkali type water electrolysis device is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell **10**.

(AEM Type Water Electrolysis Device)

[0060] Next, an example of an AEM type water electrolysis device will be described with reference to FIG. 4. As illustrated in FIG. 4, in the AEM water electrolysis device, water is supplied to the anode **13** of the electrolytic cell **10** via the anode-side water supply pipe **22**. Water passes through the membrane **11**, which is an AEM, to move from the anode **13** side to the cathode **12** side. At the cathode **12**, hydrogen and hydroxide ions (OH^-) are produced by electrolysis from water that

has passed through the membrane **11**. Hydroxide ions ($\text{OH}^{\text{sup.}-}$) produced at the cathode **12** pass through the membrane **11** to move from the cathode **12** side to the anode **13** side. At the anode **13**, oxygen and water are produced from hydroxide ions ($\text{OH}^{\text{sup.}-}$) that have passed through the membrane **11**. Water may be supplied to the cathode **12** side through the cathode-side water supply pipe **21**. Water need not be supplied through the cathode-side water supply pipe **21**.

[0061] The rate at which HDO and $\text{D}^{\text{sub.2O}}$ pass through the membrane **11** is slower than the rate at which $\text{H}^{\text{sub.2O}}$ passes through the membrane **11**. At the cathode **12**, the rate at which HD gas and $\text{D}^{\text{sub.2}}$ gas are produced from HDO and $\text{D}^{\text{sub.2O}}$ is slower than the rate at which $\text{H}^{\text{sub.2}}$ gas is produced from $\text{H}^{\text{sub.2O}}$. Thus, at the cathode **12**, the amount of deuterium gas produced, such as HD gas and $\text{D}^{\text{sub.2}}$ gas, is smaller than the amount of light hydrogen gas produced. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the AEM type water electrolysis device is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell **10**.

(Medium-High Temperature Steam Electrolysis Device)

[0062] Next, an example of a medium-high temperature steam electrolysis device will be described with reference to FIG. **5**. In the electrolysis device **1** according to the present embodiment, an electrolysis process in the electrolytic cell **10** may be SOEC (solid oxide electrolysis cell) type water electrolysis, PCEC (protonic ceramic electrolysis cell) type water electrolysis, or a combination thereof. As illustrated in FIG. **5**, the electrolysis device **1** according to the present embodiment further includes a heat exchanger **80** provided in the circulation flow path **20**. Other than this, the electrolysis device is the same as the low temperature type water electrolysis device illustrated in FIG. **1**, and therefore the description thereof will be omitted.

[0063] The heat exchanger **80** exchanges heat between water supplied to the electrolytic cell **10** and water drained from the electrolytic cell **10**. The heat exchanger **80** may include a first heat exchanger provided to extend over the cathode-side water supply pipe **21** and the cathode-side drain pipe **23**, and a second heat exchanger provided to extend over the anode-side water supply pipe **22** and the anode-side drain pipe **24**. The first heat exchanger can exchange heat between water supplied to the cathode **12** side of the electrolytic cell **10** with water discharged from the cathode **12** side of the electrolytic cell **10**. The second heat exchanger can exchange heat between water supplied to the anode **13** side of the electrolytic cell **10** with water discharged from the anode **13** side of the electrolytic cell **10**. Note that in place of the heat exchanger **80**, a heater (not illustrated) for heating water supplied to the electrolytic cell **10** may be provided in at least one of the cathode-side water supply pipe **21** or the anode-side water supply pipe **22**.

[0064] Next, electrolytic reactions in the SOEC type water electrolysis device and the PCEC type water electrolysis device will be described in detail.

(SOEC Type Water Electrolysis Device)

[0065] An example of an SOEC type water electrolysis device will be described with reference to FIG. **6**. As illustrated in FIG. **6**, in the SOEC type water electrolysis device, steam is supplied to the cathode **12** of the electrolytic cell **10** via the cathode-side water supply pipe **21**. At the cathode **12**, hydrogen gas and oxygen ions ($\text{O}^{\text{sup.2-}}$) are produced from water vapor by electrolysis. Oxygen ions ($\text{O}^{\text{sup.2-}}$) pass through the membrane **11** to move from the cathode **12** side to the anode **13** side. At the anode **13**, oxygen gas is produced from oxygen ions ($\text{O}^{\text{sup.2-}}$) that have passed through the membrane **11**.

[0066] At the cathode **12**, the rate at which HD gas and $\text{D}^{\text{sub.2}}$ gas are produced from HDO and $\text{D}^{\text{sub.2O}}$ is slower than the rate at which $\text{H}^{\text{sub.2}}$ gas is produced from $\text{H}^{\text{sub.2O}}$. Thus, at the cathode **12**, the amount of deuterium gas produced, such as HD gas and $\text{D}^{\text{sub.2}}$ gas, is smaller than the amount of light hydrogen gas produced. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the SOEC type water electrolysis device is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell **10**.

(PCEC Type Water Electrolysis Device)

[0067] Next, an example of a PCEC type water electrolysis device will be described with reference to FIG. 7. As illustrated in FIG. 7, in the PCEC type water electrolysis device, steam is supplied to the anode 13 via the anode-side water supply pipe 22. At the anode 13, oxygen gas and hydrogen ions (H.sup.+) are produced from steam by electrolysis. Hydrogen ions (H.sup.+) pass through the membrane 11 to move from the anode 13 side to the cathode 12 side. At the cathode 12, hydrogen gas is produced from hydrogen ions (H.sup.+) that have passed through the membrane 11.

[0068] The rate at which deuterium ions (D.sup.+) pass through the membrane 11 is slower than the rate at which light hydrogen ions (H.sup.+) pass through the membrane 11. At the anode 13, the rate at which D.sup.+ is produced from HDO and D.sub.2O is slower than the rate at which H.sup.+ is produced from H.sub.2O. Thus, at the cathode 12, the amount of deuterium gas produced, such as HD gas and D.sub.2 gas, is smaller than the amount of light hydrogen gas produced. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the PCEC type water electrolysis device is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell 10.

[0069] As described above, in any of the electrolysis processes, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell 10. In contrast, the abundance ratio of deuterium in water discharged from the electrolytic cell 10 is larger than the abundance ratio of deuterium in water supplied to the electrolytic cell 10. In the electrolysis device 1 according to the present embodiment, part or all of the water in the circulation flow path 20 is drained through the drainage flow path 40, and pure water is supplied to the circulation flow path 20 through the water supply flow path 30. Thus, the abundance ratio of deuterium in water flowing in the circulation flow path 20 becomes smaller by dilution, and the abundance ratio of deuterium in hydrogen molecules produced in the electrolytic cell 10 becomes even smaller. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell 10 is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0070] Note that the water supplied to the electrolytic cell 10 may be alkaline water, and the electrolysis device 1 may further include a membrane separator 90, as illustrated in FIG. 8. The membrane separator 90 may be provided in the drainage flow path 40. The membrane separator 90 may include a permeable membrane that selectively passes water in alkaline water therethrough. A semipermeable membrane selectively passes water in alkaline water therethrough. The semipermeable membrane passes water in alkaline water therethrough, but does not pass metal ions therethrough, such as sodium ions and potassium ions. Thus, only water can be discharged without discharging alkaline water to the outside of the circulation flow path 20.

[0071] The semipermeable membrane may include at least one selected from the group consisting of a flat membrane, a hollow fiber membrane, and a spiral membrane. The pore size of the semipermeable membrane may be such that water molecules pass through, but sodium ions in water to be treated do not pass through. The pore size of the semipermeable membrane may be 0.5 nm or more, or 1 nm or more. The pore size of the semipermeable membrane may be 10 nm or less, 5 nm or less, or 2 nm or less. The semipermeable membrane may be a reverse osmosis membrane (RO membrane). The semipermeable membrane may include at least one selected from the group consisting of cellulose acetate, polyacrylonitrile, polysulfone, polyethersulfone, polyethylene, polypropylene, polyvinylidene fluoride, and ceramic.

[0072] Next, the abundance ratio of deuterium to light hydrogen was evaluated by simulation under various operating conditions of the electrolysis device illustrated in FIG. 1. The operating conditions are listed in Table 1, and the abundance ratio of deuterium to light hydrogen is listed in Table 2.

TABLE-US-00001 TABLE 1 Operation conditions

1	2	3	4	5	6	7	8	9	10	11	12	Water utilization ratio				
— 0.1	0.5	Separation factor	— 600	10	600	Blow ratio	— 0.05	0.1	0.2	0.5	0.05	0.1	0.2	0.5	0.05	0.1

0.2 0.5 Amount of hydrogen Nm.sup.3/h 8000 gas generated t/d 17.1

TABLE-US-00002 TABLE 2 Operation conditions

	1	2	3	4	5	6	7	8	9	10	11	12			
Supplemental water mol-ppm	150	Electrolytic cell	mol-ppm	430	284	210	165	365	261	202	164	1542	825	448	225
supply water Hydrogen gas mol-ppm	0.7	0.5	0.3	0.3	36.5	26.1	20.2	16.4	2.57	1.38	0.75	0.38			

[0073] In Table 1, a water utilization ratio is a volume ratio of the amount of electrolysis consumption water to the amount of electrolytic cell supply water. The electrolytic cell supply water is water that is supplied to the electrolytic cell **10** through the cathode-side water supply pipe **21** and the anode-side water supply pipe **22**. The electrolysis consumption water is water to be consumed by electrolysis in the electrolytic cell **10**.

[0074] A separation factor is a value expressed by the following formula:

$$[00001] \quad = ([D]_L) / ([D]_G)$$

[0075] In the above formula, a represents the separation factor, [D].sub.L represents the abundance ratio of deuterium to light hydrogen in electrolytic cell supply water, and [D].sub.G represents the abundance ratio of deuterium to light hydrogen in hydrogen gas (hydrogen molecules) produced in the electrolytic cell **10**.

[0076] The separation factor can be controlled using an applied voltage and a current density as operating variables, which are operating conditions of the electrolytic cell **10**. The smaller the applied voltage, the smaller the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced. The smaller the current density, the smaller the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced. Furthermore, the separation factor can be controlled by selecting the metal specie of electrocatalyst. The separation factor tends to decrease in the order Cu>Fe>Ni>Ag>Au>Pt>Sn.

[0077] A blow ratio is the volume ratio of a flow rate of blow water to a flow rate of produced water. The produced water is water discharged from the electrolytic cell **10** without being electrolyzed in the electrolytic cell **10**. The blow water is water discharged from the drainage flow path **40**. The flow rate is the amount of water per unit time. The amount of hydrogen gas generated is the amount of hydrogen gas (hydrogen molecules) generated by electrolysis in the electrolytic cell **10**.

[0078] In Table 2, supplemental water is pure water supplied to the circulation flow path **20** via the water supply flow path **30**. The abundance ratio of deuterium to light hydrogen in pure water is set to 150 ppm. Electrolytic cell supply water is water supplied to the electrolytic cell **10** as described above. The electrolytic cell supply water is a mixture of supplemental water and water that circulates in the circulation flow path **20** without produced water being discharged as blow water. In this example, the abundance ratio of deuterium to light hydrogen in the electrolytic cell supply water is larger than the abundance ratio of deuterium to light hydrogen in the supplemental water.

[0079] As listed in Table 2, the abundance ratio of deuterium to light hydrogen in hydrogen gas produced in the electrolytic cell **10** can be controlled depending on operating conditions of the electrolysis device. Specifically, the lower the water utilization ratio is, the more the abundance ratio of deuterium can be reduced. The larger the separation factor is, the more the abundance ratio of deuterium can be reduced. The larger the blow ratio is, the more the abundance ratio of deuterium can be reduced.

[0080] Thus, at least one selected from the group consisting of a water utilization ratio, a separation factor, and a blow ratio may be controlled. The water utilization ratio is the ratio of the amount of water consumed by electrolysis in the electrolytic cell **10** to the amount of water supplied to the electrolytic cell. The separation factor is the ratio of the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell **10** to the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the electrolytic cell **10**. The blow ratio is the ratio of the flow rate of water discharged in the drainage step to the flow rate of water drained from the electrolytic cell **10**. By producing hydrogen molecules having a low abundance ratio of deuterium by the above-described operation, hydrogen molecules derived from fossil fuels can be more easily

distinguished. These controls may be performed by the control unit **70** by controlling the electrolytic cell **10**, the pump **25**, the pump **26**, and the flow rate control device **41**.

[0081] When the abundance ratio of deuterium in target molecules is measured, and the abundance ratio of deuterium in hydrogen molecules is within a range calculated based on operation conditions illustrated in Table 1, it can be easily determined that the target molecules are electrolytic hydrogen-containing molecules. In addition, the quality of the target molecules can be assured by checking the abundance ratio of deuterium specifications issued by manufacturers of fuels derived from renewable energy or raw materials.

[0082] As described above, the electrolysis device **1** according to the present embodiment includes the electrolytic cell **10** for electrolyzing water, the circulation flow path **20** for circulating water to be electrolyzed in the electrolytic cell **10**, and the water supply flow path **30** for supplying pure water to the circulation flow path **20**. The electrolysis device **1** includes the drainage flow path **40** for draining part or all of the water in the circulation flow path **20** downstream of the electrolytic cell **10** and upstream of the water supply via the water supply flow path **30**. The abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell **10** is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0083] The electrolysis method according to the present embodiment includes an electrolysis step for electrolyzing water in the electrolytic cell **10**, and a water supply step for supplying pure water to the circulation flow path **20**, which circulates water to be electrolyzed in the electrolytic cell **10**. The electrolysis method includes a drainage step for draining part or all of the water in the circulation flow path **20** downstream of the electrolytic cell **10** and upstream of water supply in the water supply step. The abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell **10** is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0084] In the electrolysis device and the electrolysis method according to the present embodiment, water is electrolyzed in the electrolytic cell **10**. The abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the electrolytic cell **10** is smaller than the abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell **10**. In contrast, the abundance ratio of deuterium in water discharged from the electrolytic cell **10** is larger than that of water supplied to the electrolytic cell **10**. In the electrolysis device **1** according to the present embodiment, part or all of the water in the circulation flow path **20** is drained via the drainage flow path **40**, and pure water is supplied to the circulation flow path **20** via the water supply flow path **30**. Thus, the abundance ratio of deuterium in water flowing in the circulation flow path **20** is reduced by dilution, and the abundance ratio of deuterium in hydrogen molecules produced in the electrolytic cell **10** is also reduced. Therefore, the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell **10** is smaller than the abundance ratio of deuterium to light hydrogen in nature.

[0085] In contrast, when water in the circulation flow path **20** is not drained, all of the water supplied to the electrolytic cell **10** is finally electrolyzed, and thus the abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the electrolytic cell **10** is the same as the abundance ratio of deuterium to light hydrogen in nature. Thus, the electrolysis device and the electrolysis method according to the present embodiment can reduce the abundance ratio of deuterium to light hydrogen in hydrogen molecules compared with the case where water in the circulation flow path **20** is not drained. Moreover, the abundance ratio of deuterium to light hydrogen in molecules produced can be reduced by producing molecules, such as ammonia and hydrocarbons, by using hydrogen molecules whose abundance ratio of deuterium is reduced, as a raw material.

[0086] Therefore, in the electrolysis device and the electrolysis method according to the present embodiment, it can be easily confirmed whether target molecules are hydrogen produced by water electrolysis, or molecules produced using the hydrogen as a raw material.

[0087] The entire contents of Japanese Patent Application No. 2022-175572 (filed Nov. 1, 2022) are incorporated herein by reference.

[0088] Although several embodiments have been described, modifications and variations are possible based on the above-described disclosure. All components of the above embodiments and all features described in the claims may be individually extracted and combined as long as they do not conflict with each other.

[0089] The present disclosure can contribute, for example, to goal 7 “Ensure access to affordable, reliable, sustainable and modern energy for all”, goal 12 “Ensure sustainable consumption and production patterns”, and goal 13 “Take urgent action to combat climate change and its impacts” of the Sustainable Development Goals (SDGs) led by the United Nations.

Claims

1. A determination method for determining whether or not target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules which include: hydrogen molecules produced by water electrolysis; or molecules produced using the hydrogen molecules as a raw material, the method comprising: determining that the target molecules are the electrolytic hydrogen-containing molecules when an abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than an abundance ratio of deuterium to light hydrogen in nature.
2. The determination method according to claim 1, wherein the target molecules are hydrogen molecules, ammonia, or a hydrocarbon.
3. An electrolysis device comprising: an electrolytic cell that electrolyzes water; a circulation flow path where water to be electrolyzed in the electrolytic cell circulates; a water supply flow path that supplies pure water to the circulation flow path; and a drainage flow path that drains part or all of the water in the circulation flow path, downstream of the electrolytic cell and upstream of the water supply via the water supply flow path, wherein an abundance ratio of deuterium to light hydrogen in hydrogen molecules produced by water electrolysis in the electrolytic cell is smaller than an abundance ratio of deuterium to light hydrogen in nature.
4. The electrolysis device according to claim 3, wherein a flow rate control device that controls an amount of water drained in the circulation flow path is provided in the drainage flow path.
5. The electrolysis device according to claim 3, wherein water supplied to the electrolytic cell is alkaline water, and the electrolysis device further includes a membrane separator that is provided in the drainage flow path and includes a permeable membrane which selectively passes water in the alkaline water therethrough.
6. An electrolysis method according to claim 3, comprising: an electrolysis step of electrolyzing water in the electrolytic cell; a water supply step of supplying pure water to the circulation flow path that circulates water to be electrolyzed in the electrolytic cell; and a drainage step of draining part or all of the water in the circulation flow path, downstream of the electrolytic cell and upstream of the water supply in the water supply step.
7. The electrolysis method according to claim 6, wherein at least one selected from the group consisting of: a ratio of an amount of water consumed by electrolysis in the electrolytic cell to an amount of water supplied to the electrolytic cell; a ratio of an abundance ratio of deuterium to light hydrogen in water supplied to the electrolytic cell to an abundance ratio of deuterium to light hydrogen in hydrogen molecules produced in the electrolytic cell; and a ratio of a flow rate of water drained in the drainage step to a flow rate of water discharged from the electrolytic cell, is controlled.
8. A determination device comprising: a determination unit that determines whether or not target molecules including elemental hydrogen are electrolytic hydrogen-containing molecules which include hydrogen molecules produced by water electrolysis, or molecules produced using the

hydrogen molecules as a raw material, wherein the determination unit determines that the target molecules are the electrolytic hydrogen-containing molecules when an abundance ratio of deuterium to light hydrogen in the target molecules is less than or equal to a predetermined threshold which is smaller than an abundance ratio of deuterium to light hydrogen in nature.

9. The determination device according to claim 8, further comprising: a measurement unit that measures the abundance ratio of deuterium to light hydrogen in the target molecules; a determination unit that determines that the target molecules are the electrolytic hydrogen-containing molecules when the abundance ratio of deuterium to light hydrogen in the target molecules obtained by the measurement unit is less than or equal to a predetermined threshold which is less than the abundance ratio of deuterium to light hydrogen in nature; and an output unit that outputs a determination result determined by the determination unit.
