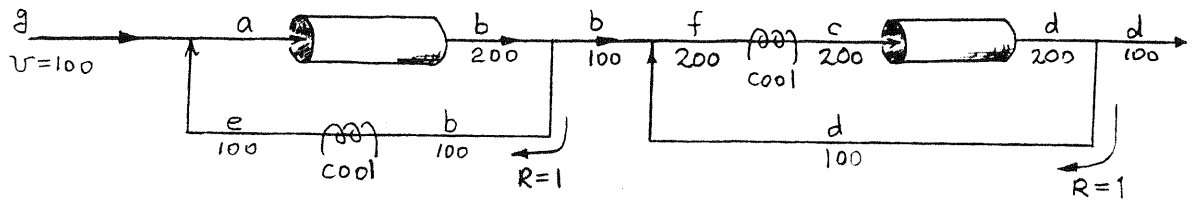
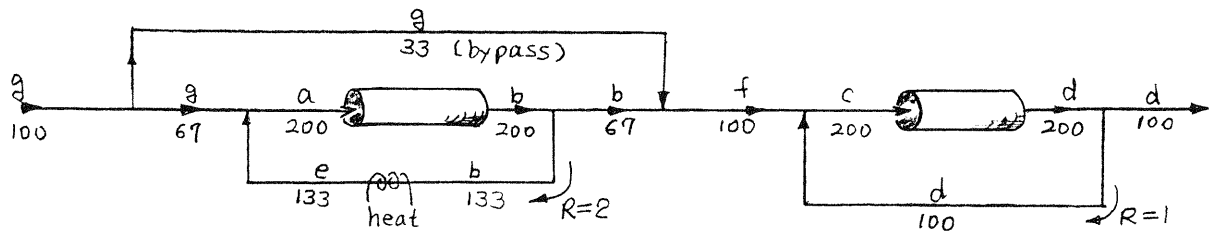


86

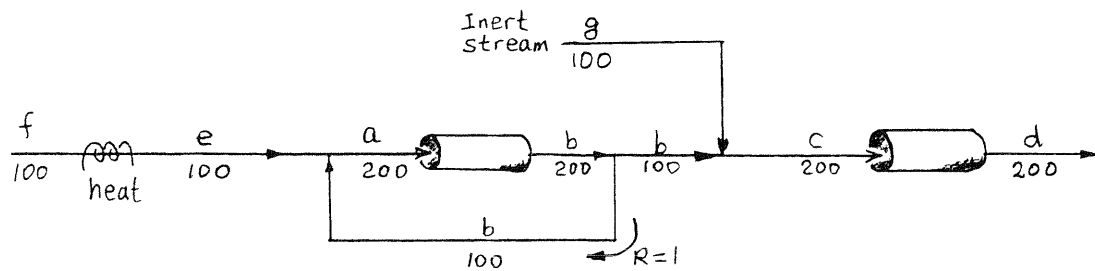
19.1



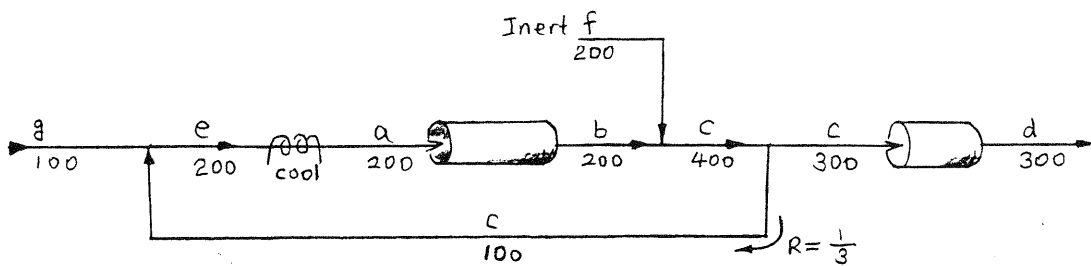
19.3



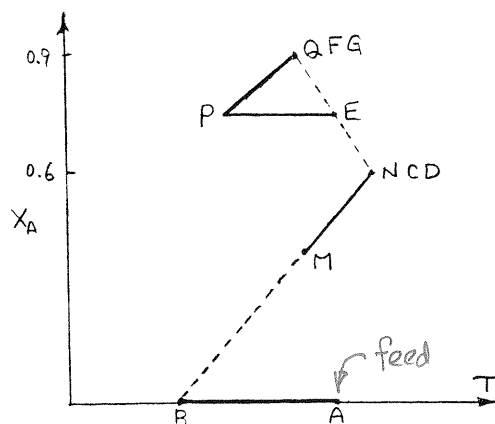
19.5



19.7

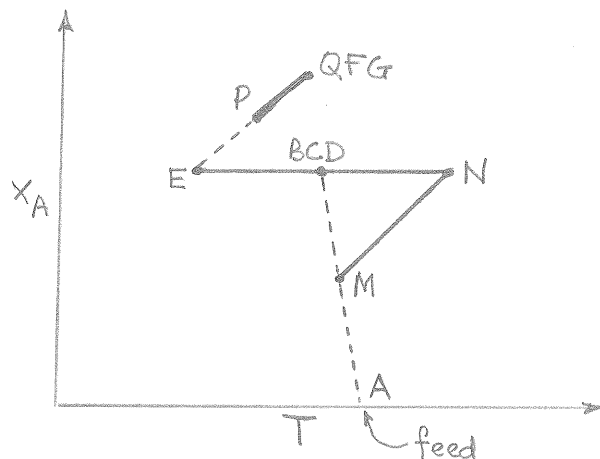


19.9



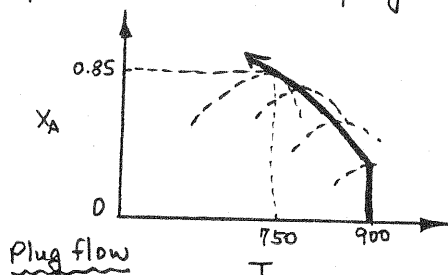
87

19.11



19.13

Optimal temperature progression



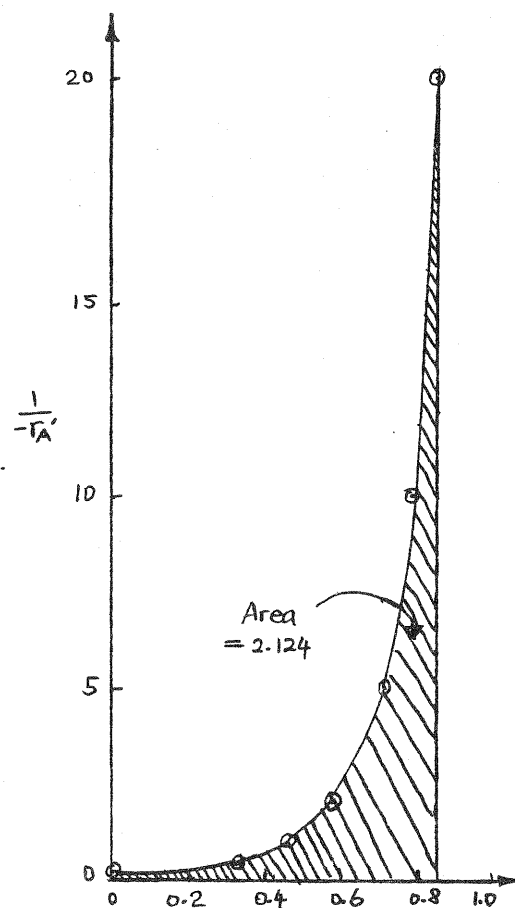
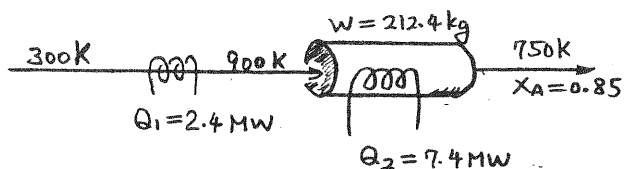
$$W = F_{A0} \int_0^{0.85} \frac{dX_A}{-r_A'} = (100)(2.124) = \underline{212.4 \text{ kg}}$$

Heat added ahead of the reactor

$$Q_1 = F_{A0} C_p \Delta T = (100)(40)(900 - 300) = \underline{2.4 \times 10^6 \text{ J/s}}$$

Heat removed from the reactor

$$\begin{aligned} Q_2 &= F_{A0} X_A (-\Delta H_r) + F_{A0} C_p (\Delta T) \\ &= (100)(0.85)(80000) + (100)(40)(900 - 750) \\ &= \underline{7.4 \times 10^6 \text{ J/s}} \end{aligned}$$



19.15 20% A - 80% inert, adiabatic reactor

$$\text{slope} = \frac{C_p}{-\Delta H_r} = \frac{40 \times 5}{80000} = \frac{1}{400}$$

Let us construct  $\frac{1}{-r_A}$  vs  $X_A$  graph.

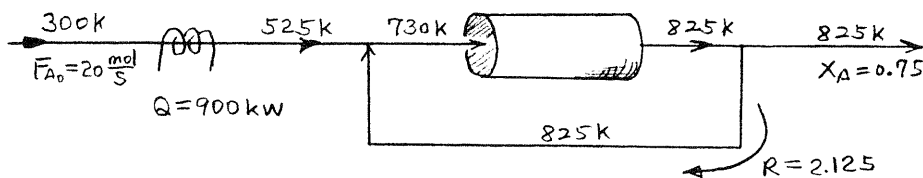
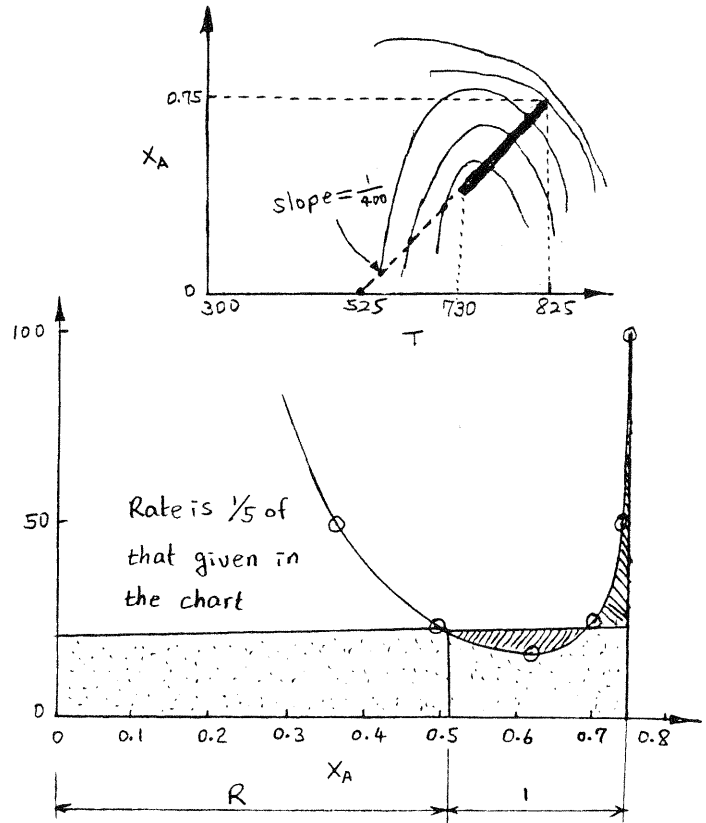
From this graph we can see that recycle flow is best and

$$R = \frac{0.51}{0.75 - 0.51} = 2.125$$

$$W = F_{A0} \frac{X_A}{-r_A'} = (100 \times 0.2)(0.75)(22.8) = 342 \text{ kg}$$

Heat duty

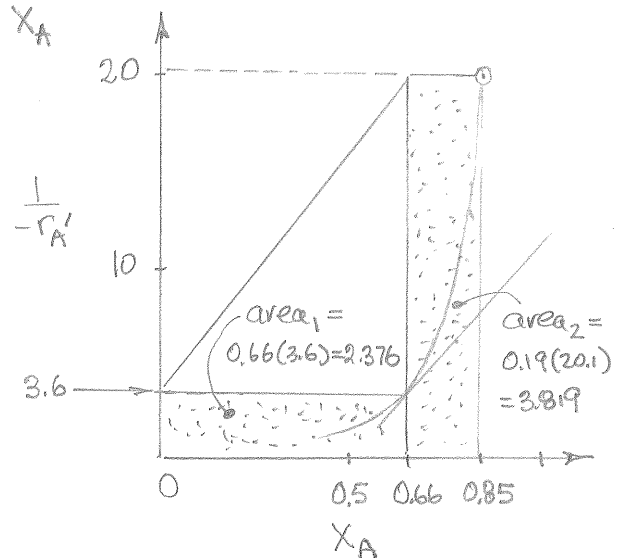
$$\begin{aligned} Q &= F_{\text{TOT}} \cdot C_p \Delta T - F_{A0} X_A (-\Delta H_r) \\ &= (100)(40)(825 - 300) - (20)(0.75)(80000) \\ &= 9 \times 10^5 \text{ J/s (heating)} \end{aligned}$$



19.17 From Fig 11 plot the  $1/(-r_{A,\text{opt}}')$  vs  $X_A$  curve. By the optimization of rectangles we find the minimum amount of catalyst needed

$$W_1 = F_{A0}(\text{area}_1) = 100(2.376) = 237.6 \text{ kg}$$

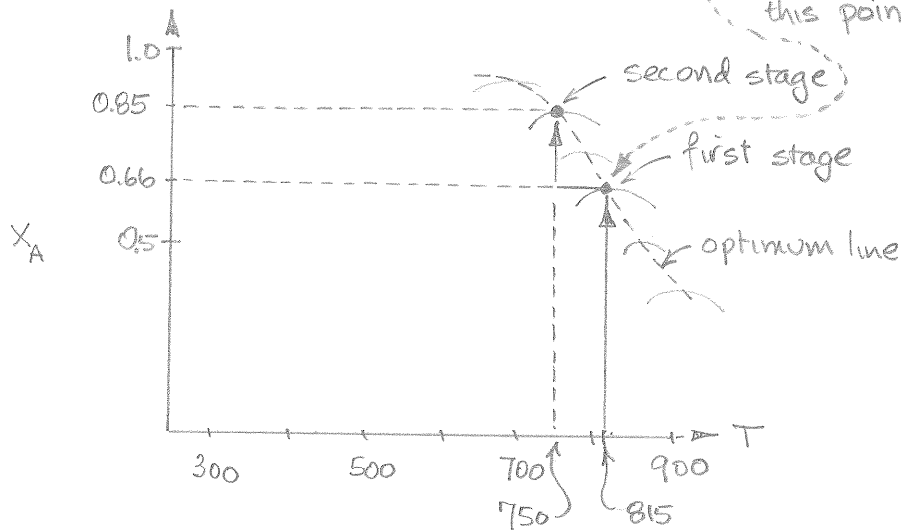
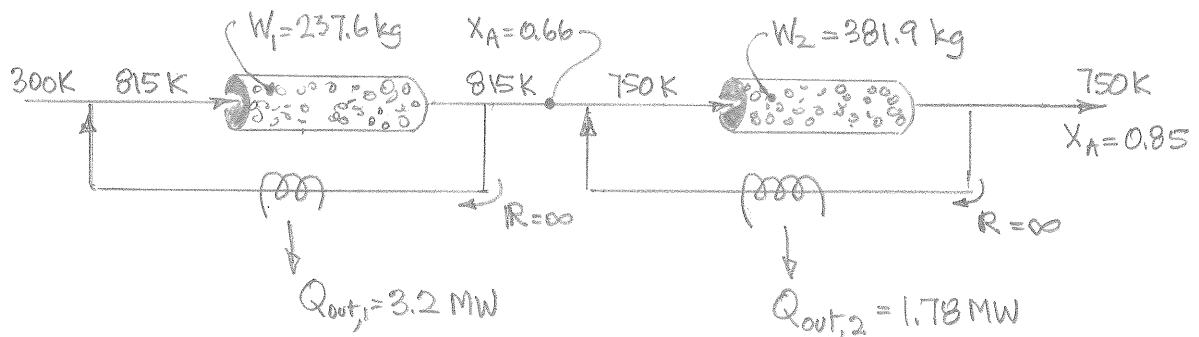
$$W_2 = F_{A0}(\text{area}_2) = 100(3.819) = 381.9 \text{ kg}$$



19.17  
(continued)Heat duty

$$\begin{aligned}
 Q_1 &= [F_{A0}(C_p \Delta T) - X_A(-\Delta H_r)] \\
 &= 100[40(815-300) - 0.66(+80000)] \\
 &= -3.2 \text{ MW (remove)}
 \end{aligned}$$

$$\begin{aligned}
 Q_2 &= [F_{A0}(C_p \Delta T) - X_A(-\Delta H_r)] \\
 &= 100[40(750-815) - 0.19(+80000)] \\
 &= -1.78 \text{ MW}
 \end{aligned}$$

The T vs  $X_A$  diagram (Fig 11) is thenFinal design

Plug flow with recycle is best.

$$W_2 = F_{A_0} \cdot \text{area}_2 = (20)(17.3) = 346 \text{ kg}$$
$$Q_1 = F_{A0} C_p (\Delta T) = (20)(200)(650 - 300) = 1.4 \times 10^6 \text{ J/s (heating)}$$

$$Q_2 = F_A C_p (\Delta T) = (20)(200)(670 - 880) \\ = -0.84 \times 10^6 \text{ J/s (cooling)}$$

Note that this solution was obtained after a few trials.

Reaction rate is  $1/5$  of that in the chart.

area 1 =  $4.64 \text{ kg.s/mol}$   
 $R_1 = 1.23$

area 2 =  $17.3 \text{ kg.s/mol}$   
 $R_2 = 0.22$

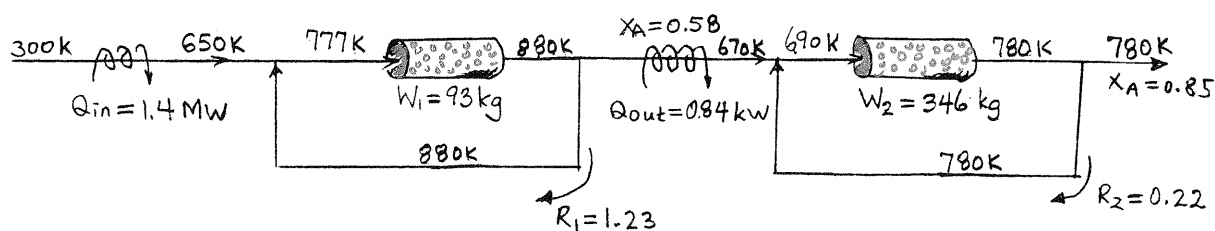
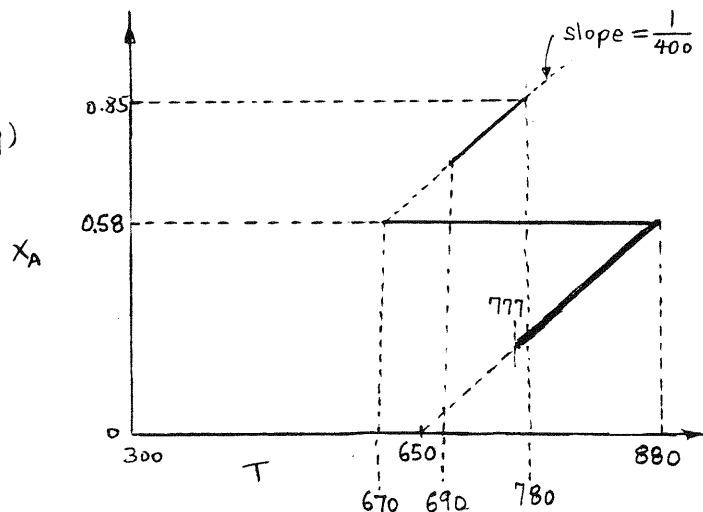
$X_A$

$R_1$   $R_2$   $I$

0 0.2 0.4 0.6 0.8

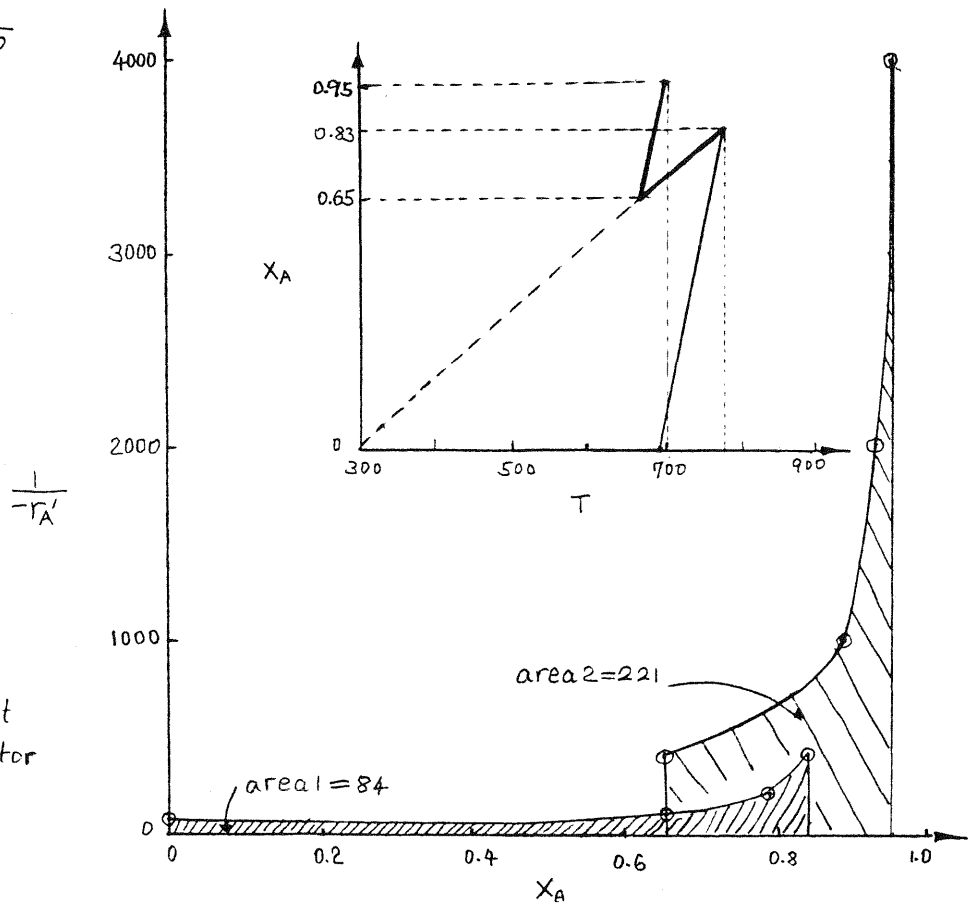
0 0.32 0.58 0.63 0.85

0 50 100 250



## 19.21 Injection of cold feed

$$\text{slope} = \frac{40 \times 20}{80000} = \frac{1}{100}$$



Fraction of feed A that goes to the first reactor

$$\frac{0.65}{0.83} = 0.7831$$

$$\begin{aligned} W_1 &= F_{A01} \cdot \text{area 1} = (5 \times 0.7831)(84) = 329 \text{ kg} \\ W_2 &= F_{A02} \cdot \text{area 2} = (5)(221) = 1105 \text{ kg} \end{aligned} \quad \left. \vphantom{\begin{aligned} W_1 &= F_{A01} \cdot \text{area 1} \\ W_2 &= F_{A02} \cdot \text{area 2} \end{aligned}} \right\} W_{\text{total}} = 1434 \text{ kg}$$

$$Q = F_{A01} \cdot C_p (\Delta T) = (5 \times 0.7831)(800)(670 - 300) = 1.22 \times 10^6 \text{ J/s (heating)}$$

Note: only one heat exchanger is needed in this design.

