Of course, I'd be happy to help! The discovery of atoms can be traced back to the ancient Greek s, but it was not until the 19th century that the concept of atoms as we know it today was develo ped. Here is a brief summary:

* Ancient Greece (5th - 4th century BCE): The Greek philosopher Democritus proposed the idea of tiny indivisible particles called "atomos" that made up all matter. He believed that atoms were eternal and could not be created or destroyed.
* Dalton's Atomic Theory (1803): English chemist John Dalton proposed the modern atomic the ory, which states that elements are composed of small, indivisible particles called atoms, and th at compounds are formed when atoms of different elements combine in whole number ratios.
* Thomson's Discovery (1897): Scottish physicist J.J. Thomson discovered the electron, which l ed to the understanding that atoms also contain negatively charged electrons. This discovery co nfirmed that atoms are made up of smaller particles than previously thought.
* Rutherford's Nuclear Model (1911): Ernest Rutherford conducted experiments that led him to p ropose the nuclear model of the atom, where most of an atom's mass is concentrated in a small, dense nucleus at the center of the atom. This model replaced the earlier plum pudding model o f the atom.
* Bohr's Quantization Hypothesis (1913): Niels Bohr suggested that energy levels in atoms are quantized, meaning they only allow certain specific energies for electrons to occupy. This hypot hesis helped explain why some chemical reactions occur more easily than others.

These discoveries and theories laid the foundation for our current understanding of atoms as th e building blocks of matter, and have had a profound impact on our understanding of the univer se and the laws of physics.