Of course, I'd be happy to help! The discovery of atoms dates back to the ancient Greeks, but it was not until the 19th century that the concept of atoms as we know it today was developed. He re is a brief summary:

* Ancient Greece: The idea of indivisible particles that made up matter can be traced back to the ancient Greeks. Philosophers such as Democritus proposed that matter was composed of tiny i ndestructible particles called "atomos," which could not be created or destroyed, only rearrange d.
* John Dalton: In the early 19th century, English chemist and physicist John Dalton proposed th e modern atomic theory, which states that elements are composed of small, indivisible particles called atoms, and that compounds are formed when atoms of different elements combine in wh ole number ratios. He also introduced the concept of the atom as having a specific mass and siz e.
* J.J. Thomson: In 1897, British physicist J.J. Thomson discovered the electron, which he believ ed to be a fundamental particle smaller than an atom. This led to the realization that atoms were composed of even smaller particles, such as electrons, protons, and neutrons.
* Ernest Rutherford: Canadian physicist Ernest Rutherford conducted experiments in 1911 that showed that atoms had a nucleus at their center, surrounded by electrons. His experiment invol ved bombarding atoms with alpha particles (now known as helium nuclei) and observing the sca ttering pattern produced by the collisions. This led to the discovery of the nucleus and the devel opment of the nuclear model of the atom.
* Niels Bohr: Danish physicist Niels Bohr further developed the atomic model in the early 20th c entury. He proposed that the electrons in an atom occupy specific energy levels around the nucl eus, and that the atom could only exist in certain discrete states. His model explained many obs erved properties of atoms, such as the emission spectra of elements and the stability of atoms.

These discoveries and others have contributed to our understanding of the structure of atoms a nd the way they interact with each other to form the world around us.