Of course, I'd be happy to help! The discovery of atoms dates back to the ancient Greeks, but it wasn't until the 19th century that the concept of atoms as we know it today was developed. Her e's a brief summary:

* Ancient Greece (5th - 4th century BCE): The Greek philosopher Democritus proposed that mat ter was composed of tiny indivisible particles called "atomos," which could not be created or des troyed, only rearranged. This idea laid the foundation for the concept of atoms.
* John Dalton (1803): English chemist and physicist John Dalton proposed the modern atomic th eory, which states that elements are composed of small, indivisible particles called atoms, and t hat compounds are formed when atoms of different elements combine in whole number ratios. He also introduced the concept of the atomic model, where atoms were seen as tiny balls that m oved independently in space.
* J.J. Thomson (1897): British physicist J.J. Thomson discovered the electron, which is a subato mic particle with a negative charge. His work showed that atoms did not have an infinite number of electrons, as previously thought, but rather had a fixed number of electrons.
* Ernest Rutherford (1911): New Zealand-born physicist Ernest Rutherford conducted experime nts that led him to propose the nuclear model of the atom, where the positive charge of an atom resides within a small, dense nucleus surrounded by electrons. This model revolutionized our u nderstanding of the structure of atoms.
* Niels Bohr (1913): Danish physicist Niels Bohr further developed the atomic model by proposin g the quantum mechanical nature of the atom. According to his model, electrons occupy specific energy levels around the nucleus, and can jump between these levels through the emission or absorption of energy in the form of photons.

These discoveries and theories formed the basis of modern atomic physics and continue to sha pe our understanding of the universe at its most fundamental level.