

US Patent & Trademark Office

Patent Public Search | Text View

United States Patent Application Publication

20250266445

Kind Code

A1

Publication Date

August 21, 2025

Inventor(s)

DENG; Haixia et al.

Materials and Methods of Producing Lithium Cobalt Oxide Materials of A Battery Cell

Abstract

Various lithium cobalt oxides materials doped with one or more metal dopants having a chemical formula of $\text{Li.sub.xCo.sub.yO.sub.z}$, and method and apparatus of producing the various lithium cobalt oxides materials are provided. The method includes adjusting a molar ratio $\text{M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt:} \dots \text{M.sub.MeNSalt}$ of a lithium-containing salt, a cobalt-containing salt and one or more metal-dopant-containing salts within a liquid mixture to be equivalent to a ratio of $x:y:a:b:c: \dots n$, drying a mist of the liquid mixture in the presence of a gas to form a gas-solid mixture, separating the gas-solid mixture into one or more solid particles of an oxide material, and annealing the solid particles of the oxide material in the presence of another gas flow to obtain crystalized particles of the lithium cobalt oxide material.

Inventors: DENG; Haixia (Fremont, CA), LIU; Shengfeng (Newark, CA), LIU; Min-Duan (Bethany, CT), LIU; Mengchen (Union City, CA), CHEN; Liang-Yuh (Saratoga, CA)

Applicant: eJoule, Inc. (Fremont, CA)

Family ID: 1000008576740

Appl. No.: 19/186616

Filed: April 23, 2025

Related U.S. Application Data

parent US continuation 17478869 20210917 PENDING child US 19186616
us-provisional-application US 63080023 20200918

Publication Classification

Int. Cl.: H01M4/525 (20100101); C01G51/42 (20250101); H01M4/02 (20060101); H01M4/04 (20060101)

U.S. Cl.:

CPC H01M4/525 (20130101); C01G51/42 (20130101); H01M4/0471 (20130101); C01P2002/72 (20130101); C01P2004/03 (20130101); C01P2006/40 (20130101); H01M2004/028 (20130101)

Background/Summary

CROSS-REFERENCE TO RELATED APPLICATIONS [0001] This application is a continuation of U.S. patent application Ser. No. 17/478,869, filed on Sep. 17, 2021, which claims benefit of U.S. provisional patent application Ser. No. 63/080,023, filed on Sep. 18, 2020, which. All of the above-referenced applications are herein incorporated by reference.

BACKGROUND OF THE INVENTION

[0002] Great efforts have been devoted to the development of advanced electrochemical battery cells to meet the growing demand of various consumer electronics, electrical vehicles and grid energy storage applications in terms of high energy density, high power performance, high capacity, long cycle life, low cost and excellent safety. Thus, a need for more efficient utilization of the available energy resources as well as air-quality-control has generated an enormous interest in the development of advanced high energy density batteries for electric powered vehicles. Furthermore, cost effectiveness, great cycling life, stability, rechargeability, and better safety characteristics have been other factors driving the battery market.

[0003] In an electrochemically active battery cell, a cathode and an anode are immersed in an electrolyte and electronically separated by a separator. The separator is typically made of porous polymer membrane materials such that metal ions released from the electrodes into the electrolyte can diffuse through the pores of the separator and migrate between the cathode and the anode during battery charge and discharge. The type of a battery cell is usually named from the metal ions that are transported between its cathode and anode electrodes. Lithium ion battery is a secondary battery which was developed in the early 1990s and it represent a new generation of lightweight, compact, and yet high-energy power sources. However, the cost for commercially manufacturing various lithium battery materials is considerably higher than other types of secondary batteries.

[0004] Cathode active materials are the most expensive component in a lithium ion battery and, to a relatively large extent, determines the energy density, cycle life, manufacturing cost and safety of a lithium battery cell. Examples of good cathode active materials include nanometer- or micron-sized lithium transition metal oxide materials and lithium ion phosphate, etc. When lithium battery was first commercialized, lithium cobalt oxide (LiCoO_2) material is used as the cathode material. While the theoretical capacity of LiCoO_2 is about 274-275 mAh/g, and a capacity of the LiCoO_2 when using 4.2 V as an upper limit voltage is about 150 mAh/g.

[0005] To further increase the battery performance of LiCoO_2 , one can increase charging cut-off voltage to extract more Li^{+} . However, conventional material manufacturing processes such as solid-state reaction (e.g., mixing solid precursors and then calcination) and wet-chemistry processes (e.g., treating precursors in solution through co-precipitation, sol-gel, or hydrothermal reaction, etc., and then mixing and calcination) have notable challenges in promoting cycle stability of LiCoO_2 at high voltage. Since a high voltage is applied to LiCoO_2 materials, it is difficult to consistently produce LiCoO_2 having the characteristics of high stability and long battery life cycle at a level of industrial size.

[0006] In addition, solid-state diffusion rates affect the performance of resulting batteries made from these lithium oxide materials in applications requiring high-powered batteries. Overall, the processing time for such a solid-state multi-step batch manufacturing process will take up to a week so it is very labor intensive and energy consuming. Batch process also increases the chance of introducing impurity with poor run-to-run quality consistency and low overall yield. Specifically, co-precipitation is not suitable for the preparation of highly pure, accurate stoichiometric phases of these lithium-containing transition metal oxide battery materials.

[0007] Thus, there is a need for an improved method and system to manufacture high power performance, high capacity, long cycle life, excellent stability, properly crystalized, structured lithium metal oxide active materials for a lithium-ion battery (LIB) cell at high voltage and high temperature.

SUMMARY OF THE INVENTION

[0008] Embodiments of the invention generally provide lithium ion battery materials and methods for producing lithium ion battery materials thereof. One embodiment of the invention provides an oxide material, such as a lithium cobalt oxide material doped with at least one metal dopant ($\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}\text{Math.doped Me}_{\text{sub.a}}$), wherein x is from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and $0 < a \leq 0.05$. The material can be obtained from a process, which includes forming a mist of a liquid mixture comprising a lithium-containing salt, a cobalt-containing salt, and at least one metal-dopant-containing salt, mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture, wherein the liquid mixture is soluble in a suitable solvent, drying the gas-liquid mixture to form a gas-solid mixture, separating the gas-solid mixture into one or more solid particles of an oxide material; and annealing the one or more solid particles of the oxide material at an annealing temperature of 400°C . or higher to obtain crystalized particles of the lithium cobalt oxide material doped with at least one metal dopant ($\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}\text{Math.doped Me}_{\text{sub.a}}$).

[0009] In one example, the lithium cobalt oxide material doped with the at least one metal dopant ($\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}\text{Math.doped Me}_{\text{sub.a}}$), is obtained from adjusting a molar ratio $\text{M}_{\text{sub.LiSalt}}:\text{M}_{\text{sub.CoSalt}}:\text{M}_{\text{sub.MeSalt}}$ of the lithium-containing salt, the cobalt-containing salt, and the at least one metal-dopant-containing salts in the liquid mixture to be a ratio of about $x:y:a$ for making the lithium cobalt oxide material doped with at least one metal dopant (Me) at desirable atomic ratio of $\text{Li}:\text{Co}:\text{Me}$ equaling to $x:y:a$. For example, the molar ratio $\text{M}_{\text{sub.LiSalt}}:\text{M}_{\text{sub.CoSalt}}:\text{M}_{\text{sub.MeSalt}}$ of the lithium-containing salt, the cobalt-containing salt, and the at least one metal-dopant-containing salt is performed prior to forming the mist of the liquid mixture. As another example, molar ratio $\text{M}_{\text{sub.LiSalt}}:\text{M}_{\text{sub.CoSalt}}:\text{M}_{\text{sub.MeSalt}}$ of the lithium-containing salt, the cobalt-containing salt, and the at least one metal-dopant-containing salts can be adjusted at the same time of forming the mist of the liquid mixture.

[0010] Another embodiment of the invention provides a lithium cobalt oxide material doped with one or more metal dopants ($\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}\text{Math.doped Me1}_{\text{sub.a}}, \text{Me2}_{\text{sub.b}}, \text{Me3}_{\text{sub.c}}, \dots \text{MeN}_{\text{sub.n}}$), wherein x is from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and wherein $N \geq 1$, and each a, b, c, \dots, n is more than 0 and no more than 0.05. The lithium cobalt oxide material doped with one or more metal dopants is obtained from a process, which includes adjusting a molar ratio $\text{M}_{\text{sub.LiSalt}}:\text{M}_{\text{sub.CoSalt}}:\text{M}_{\text{sub.Me1Salt}}:\text{M}_{\text{sub.Me2Salt}}:\text{M}_{\text{sub.Me3Salt}}:\dots \text{M}_{\text{sub.MeNSalt}}$ of a lithium-containing salt, a cobalt-containing salt, and one or more metal-dopant-containing salts which are soluble in a suitable solvent into a liquid mixture, wherein each of the one or more metal-dopant-containing salts is selected from a group consisting of a first metal-containing salt, a second metal-containing salt, a third metal-containing salt, \dots an N metal-containing salt and combinations thereof, and forming a mist of the liquid mixture. The process further includes mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture, drying the gas-liquid mixture to form a gas-solid mixture, separating the gas-solid mixture into one or more solid

particles of an oxide material; and annealing the one or more solid particles of the oxide material at an annealing temperature of 400° C. or higher to obtain crystalized particles of the lithium cobalt oxide material doped with one or more metal dopants.

[0011] Another embodiment of the invention provides a lithium cobalt oxide material doped with at least one metal dopant ($\text{Li}_{0.9-1.1}\text{Co}_{0.9-1.1}\text{O}_{1.8-2.2}\text{Me}_a$), wherein x is from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and wherein $0 < a \leq 0.05$. The lithium cobalt oxide material doped with one or more metal dopants is obtained from a process, which includes adjusting a molar ratio $\text{M}_{\text{LiSalt}}:\text{M}_{\text{CoSalt}}:\text{M}_{\text{MeSalt}}$ of a lithium-containing salt, a cobalt-containing salt, and at least one metal-dopant-containing salt into a liquid mixture and forming a mist of the liquid mixture, where the liquid mixture comprises the lithium-containing salt, the cobalt-containing salt; the at least one metal-dopant-containing salt; and a suitable solvent. The process further includes mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture, drying the gas-liquid mixture to form one or more solid particles of an oxide material; and annealing the one or more solid particles of the oxide material at an annealing temperature of 400° C. or higher to obtain crystalized particles of the lithium cobalt oxide material doped with at least one metal dopant ($\text{Li}_{0.9-1.1}\text{Co}_{0.9-1.1}\text{O}_{1.8-2.2}\text{Me}_a$).

[0012] In yet another embodiment, a method of producing a lithium cobalt oxide material with one or more metal dopants having a chemical formula of $\text{Li}_{0.9-1.1}\text{Co}_{0.9-1.1}\text{O}_{1.8-2.2}\text{Me}_a$ ($\text{Me}_1, \text{Me}_2, \text{Me}_3, \dots, \text{Me}_n$) is provided. The method includes forming a mist of a liquid mixture, where the liquid mixture is obtained from adjusting a molar ratio $\text{M}_{\text{LiSalt}}:\text{M}_{\text{CoSalt}}:\text{M}_{\text{MeSalt}}$ of a lithium-containing salt (LiSalt), a cobalt-containing salt (CoSalt), and at least one metal-dopant-containing salt (MeSalt) in the liquid mixture to be a ratio of about x:y:a for making the lithium cobalt oxide material doped with at least one metal dopant (Me) at desirable atomic ratio of Li:Co:Me equaling to x:y:a. The method further includes mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture, drying the gas-liquid mixture to form a gas-solid mixture, separating the gas-solid mixture into one or more solid particles of an oxide material, and annealing the one or more solid particles of the oxide material at an annealing temperature of 400° C. or higher to obtain crystalized particles of the lithium cobalt oxide material doped with at least one metal dopant (Me) ($\text{Li}_{0.9-1.1}\text{Co}_{0.9-1.1}\text{O}_{1.8-2.2}\text{Me}_a$), where x is from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and $0 < a \leq 0.05$.

Description

BRIEF DESCRIPTION OF THE DRAWING

[0013] So that the manner in which the above recited features of the present invention can be understood in detail, a more particular description of the invention, briefly summarized above, may be had by reference to embodiments, some of which are illustrated in the appended drawings. It is to be noted, however, that the appended drawings illustrate only typical embodiments of this invention and are therefore not to be considered limiting of its scope, for the invention may admit to other equally effective embodiments.

[0014] FIG. 1A illustrates one embodiment of a flow chart of a method of producing cathode materials for lithium ion batteries.

[0015] FIG. 1B illustrates another embodiment of another flow chart of a method of producing cathode materials for lithium ion batteries.

[0016] FIG. 2 is a schematic of an exemplary processing system useful in preparing a material for a battery electrochemical cell according one embodiment of the invention.

[0017] FIG. 3 is a schematic of another exemplary processing system useful in preparing a material

for a battery electrochemical cell according one embodiment of the invention.

[0018] FIG. 4 is a line graph illustrating the discharge profile of electric capacity of lithium ion batteries prepared from various cathode materials of the invention.

[0019] FIG. 5A is a column graph illustrating the discharge profile of electric capacity of lithium ion batteries at a specified voltage where the lithium ion batteries are prepared from various exemplary cathode materials of the invention.

[0020] FIG. 5B is a column graph illustrating the discharge profile of electric capacity of lithium ion batteries at a specified voltage where the lithium ion batteries are prepared from exemplary cathode materials of the invention.

[0021] FIG. 5C is a column graph illustrating the discharge profile of electric capacity of lithium ion batteries at a specified voltage where the lithium ion batteries are prepared from yet several examples of a cathode material of the invention.

[0022] FIG. 6 is a graph illustrating electric charge and discharge cycling performance of battery cells prepared by using various examples of cathode materials of the invention.

[0023] FIG. 7A is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0024] FIG. 7B is a scanning electron microscopy (SEM) image of the example of FIG. 7A in larger magnitude.

[0025] FIG. 7C is a scanning electron microscopy (SEM) image of another embodiment of another example of solid particles of an oxide material after a drying process.

[0026] FIG. 7D is a scanning electron microscopy (SEM) image of the example of FIG. 7C in larger magnitude.

[0027] FIG. 8A is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials.

[0028] FIG. 8B is a scanning electron microscopy (SEM) image of the example of FIG. 8A in larger magnitude.

[0029] FIG. 8C is a scanning electron microscopy (SEM) image of another embodiment of another example of solid particles of an oxide material after a drying process.

[0030] FIG. 8D is a scanning electron microscopy (SEM) image of the example of FIG. 8C in larger magnitude.

[0031] FIG. 9A is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0032] FIG. 9B is a scanning electron microscopy (SEM) image of the example of FIG. 9A in larger magnitude.

[0033] FIG. 9C is a scanning electron microscopy (SEM) image of another embodiment of another example of solid particles of an oxide material after a drying process.

[0034] FIG. 9D is a scanning electron microscopy (SEM) image of the example of FIG. 9C in larger magnitude.

[0035] FIG. 10A is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0036] FIG. 10B is a scanning electron microscopy (SEM) image of the example of FIG. 10A in larger magnitude.

[0037] FIG. 10C is a scanning electron microscopy (SEM) image of another embodiment of another example of solid particles of an oxide material after a drying process.

[0038] FIG. 10D is a scanning electron microscopy (SEM) image of the example of FIG. 10C in larger magnitude.

[0039] FIG. 11A is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0040] FIG. 11B is a scanning electron microscopy (SEM) image of the example of FIG. 11A in larger magnitude.

[0041] FIG. **12A** is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0042] FIG. **12B** is a scanning electron microscopy (SEM) image of the example of FIG. **12A** in larger magnitude.

[0043] FIG. **13A** is a scanning electron microscopy (SEM) image of another embodiment of another example of solid particles of an oxide material after a drying process.

[0044] FIG. **13B** is a scanning electron microscopy (SEM) image of the example of FIG. **13A** in larger magnitude.

[0045] FIG. **14A** is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0046] FIG. **14B** is a scanning electron microscopy (SEM) image of the example of FIG. **14A** in larger magnitude.

[0047] FIG. **14C** is a scanning electron microscopy (SEM) image of the example of FIG. **14A** in larger magnitude.

[0048] FIG. **15A** is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0049] FIG. **15B** is a scanning electron microscopy (SEM) image of the example of FIG. **15A** in larger magnitude.

[0050] FIG. **15C** is a scanning electron microscopy (SEM) image of the example of FIG. **15A** in larger magnitude.

[0051] FIG. **16A** is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0052] FIG. **16B** is a scanning electron microscopy (SEM) image of the example of FIG. **16A** in larger magnitude.

[0053] FIG. **16C** is a scanning electron microscopy (SEM) image of the example of FIG. **16A** in larger magnitude.

[0054] FIG. **17A** is a scanning electron microscopy (SEM) image of another example of crystalized lithium cobalt oxide materials of the invention.

[0055] FIG. **17B** is a scanning electron microscopy (SEM) image of the example of FIG. **17A** in larger magnitude.

[0056] FIG. **17C** is a scanning electron microscopy (SEM) image of the example of FIG. **17A** in larger magnitude.

[0057] FIG. **18** is an X-ray diffraction (XRD) pattern of two examples of crystalized lithium cobalt oxide materials of the invention.

DETAILED DESCRIPTION

[0058] This invention generally relates to compositions, oxide materials, battery materials, apparatuses, and methods thereof in soluble solutions in proper molar ratio to precisely control and obtain proper atomic-level ratios and make-up of a battery active material to be used for a lithium-ion battery. The battery materials and methods and apparatus provided here results in highly pure, accurate stoichiometric phases battery cathode materials and can be used, in turn, to make lithium-ion batteries with, with characteristics associated with high battery cycling performance, including high electric capacity.

[0059] FIG. **1A** is a flow chart showing a method **100** of producing lithium cobalt oxide material doped with one or more metal dopants having a chemical formula of

$\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}\text{Math.doped Me1}_{\text{sub.a}}\text{Me2}_{\text{sub.b}}\text{Me3}_{\text{sub.c}} \dots \text{MeN}_{\text{sub.n}}$ for lithium-ion batteries. The method **100** includes a step **110** or series of steps of adjusting a molar ratio $\text{M}_{\text{sub.LiSalt}}:\text{M}_{\text{sub.CoSalt}}:\text{M}_{\text{sub.Me1Salt}}:\text{M}_{\text{sub.Me2Salt}}:\text{M}_{\text{sub.Me3Salt}}: \dots :\text{M}_{\text{sub.MeNSalt}}$ of a lithium-containing salt (LiSalt), a cobalt-containing salt (CoSalt), and one or more metal-dopant-containing salts which are soluble in a suitable solvent into a liquid mixture, where each of the one or more metal-dopant-containing salts is selected from a group consisting of a first metal-

containing salt, a second metal-containing salt, a third metal-containing salt, . . . an N metal-containing salt and combinations thereof. The molar ratio $M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:M_{\text{sub.Me3Salt}}:\dots:M_{\text{sub.MeNSalt}}$ of the lithium-containing salt (LiSalt), the cobalt-containing salt (CoSalt), and the one or more metal-dopant-containing salts is adjusted to be a ratio of about $x:y:a:b:c:\dots:n$ for making the lithium cobalt oxide doped with one or more metal dopants ($\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}$ Math.doped $\text{Me1}_{\text{sub.a}}\text{Me2}_{\text{sub.b}}\text{Me3}_{\text{sub.c}}\dots\text{MeN}_{\text{sub.n}}$) at desirable atomic ratio of Li:Co:Me1:Me2:Me3 . . . :MeN equaling to $x:y:a:b:c:\dots:n$, where x is from 0.9 to 1.1 ($0.9\leq x\leq 1.1$), y is from 0.9 to 1.1 ($0.9\leq y\leq 1.1$), z is from 1.8 to 2.2 ($1.8\leq z\leq 2.2$), and where $N\geq 1$, a is from 0 to 0.05 ($0\leq a\leq 0.05$), b is from 0 to 0.05 ($0\leq b\leq 0.05$), c is from 0 to 0.05 ($0\leq c\leq 0.05$), . . . , and n is from 0 to 0.05 ($0\leq n\leq 0.05$).

[0060] In one embodiment, the desired molar ratio of

$M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:M_{\text{sub.Me3Salt}}:\dots:M_{\text{sub.MeNSalt}}$ can be achieved by measuring and preparing appropriate amounts a lithium-containing salt (LiSalt), a cobalt-containing salt (CoSalt), a first metal dopant-containing salt (Me1Salt), a second metal dopant-containing salt (Me2Salt), a third metal dopant-containing salt (Me3Salt), . . . , and a N metal dopant-containing salt (MeNSalt). For example, the molar ratio

$M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:M_{\text{sub.Me3Salt}}:\dots:M_{\text{sub.MeNSalt}}$ of the lithium-containing salt, the cobalt-containing salt, the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . the N metal-containing salt can be adjusted (e.g., manually or digitally using a processing system of the invention) and prepared directly into a liquid mixture in a desired concentration prior to forming the mist of the liquid mixture. As another example, the adjusting the molar ratio

$M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:M_{\text{sub.Me3Salt}}:\dots:M_{\text{sub.MeNSalt}}$ of the lithium-containing salt, the cobalt-containing salt, the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . and the N-containing salt can be performed simultaneously with forming the mist of the liquid mixture.

[0061] The method **100** includes further includes a step **120** of forming a liquid mixture having the lithium-containing salt at the molarity of $M_{\text{sub.LiSalt}}$, the cobalt-containing salt at the molarity of $M_{\text{sub.CoSalt}}$, and the one or more metal dopant-containing salts (e.g., a First Metal-Containing Salt at a Molarity of $M_{\text{sub.Me1Salt}}$, a Second Metal-Containing Salt at a Molarity of $M_{\text{sub.Me2Salt}}$, an N Metal-Containing Salt at a Molarity of . . . $M_{\text{sub.MeNSalt}}$, etc.) for producing lithium cobalt oxide materials doped with one or more metal dopants with a targeting formula of $\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}$ Math.doped $\text{Me1}_{\text{sub.a}}\text{Me2}_{\text{sub.b}}\text{Me3}_{\text{sub.c}}\dots\text{MeN}_{\text{sub.n}}$, where the one or more metal-dopant-containing salts comprising the first-containing metal salt, the second-containing metal salt, the third-containing metal salt, . . . , the N-containing salt are generated, and where the liquid mixture achieves the molar ratio of

$M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:\dots M_{\text{sub.MeNSalt}}$ at about of $x:y:a:b:\dots:n$.

[0062] The mist of the liquid mixture may include droplets of various reactant solution, precursor solutions, etc., in homogenous forms, sizes, shape, etc. For example, the molar ratio

$M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:M_{\text{sub.Me3Salt}}:\dots:M_{\text{sub.MeNSalt}}$ of the lithium-containing salt, the cobalt-containing salt, and the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . the N-containing salt can be digitally adjusted, depending on the desired composition of final solid product particles.

[0063] In one embodiment, the one or more metal dopants (Me1, Me2, Me3, . . . MeN) are incorporated into the lithium cobalt oxide materials, wherein Me1, Me2, Me3, . . . MeN are different metal dopants. For example, each of the one or more metal dopants (i.e. Me1, Me2, Me3, . . . MeN) can be selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

[0064] Another embodiment of the present invention is that the liquid form of the lithium-containing salt, the cobalt-containing salt and the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . the N-containing salt can be dissolved or dispersed in a suitable solvent (e.g., water, alcohol, methanol, isopropyl alcohol, organic solvents, inorganic solvents, organic acids, sulfuric acid (H_2SO_4), citric acid ($\text{C}_6\text{H}_8\text{O}_7$), acetic acids (CH_3COOH), butyric acid ($\text{C}_4\text{H}_8\text{O}_2$), lactic acid ($\text{C}_3\text{H}_6\text{O}_3$), nitric acid (HNO_3), hydrochloric acid (HCl), ethanol, pyridine, ammonia, acetone, and their combinations) to form into a liquid mixture of an aqueous solution, slurry, gel, aerosol or any other suitable liquid forms. For example, one or more solid particles of an oxide material can be adjusted manually or digitally and prepared in desirable molar ratio and mixed into a liquid mixture, such as by adjusting, measuring and preparing appropriate amounts of the lithium-containing salt compound, the cobalt-containing salt compound and the one or more metal-dopant-containing salts into one solution with suitable amounts of a solvent. Depending on the solubility of the lithium-containing salt, the cobalt-containing salt and the one or more metal-containing salts in a chosen solvent, pH, temperature, and mechanical stirring and mixing can be adjusted to obtain a liquid mixture where the one or more metal-dopant-containing salts at the desirable molar concentrations are fully dissolved and/or evenly dispersed.

[0065] In another embodiment, the lithium containing salts are mixed into the liquid mixture. Exemplary lithium containing salts include, but not limited to, lithium sulfate (Li_2SO_4), lithium nitrate (LiNO_3), lithium carbonate (Li_2CO_3), lithium acetate (LiCH_3COO), lithium hydroxide (LiOH), lithium formate (LiCHO_2), lithium chloride (LiCl), and combinations thereof. The cobalt containing salts are mixed into the liquid mixture. Exemplary cobalt containing salts include, but not limited to, cobalt sulfate (CoSO_4), cobalt nitrate ($\text{Co}(\text{NO}_3)_2$), cobalt acetate ($\text{Co}(\text{CH}_3\text{COO})_2$), cobalt formate ($\text{Co}(\text{CHO}_2)_2$), cobalt chloride (CoCl_2), and combinations thereof.

[0066] In still another embodiment, the first metal-containing salt, the second metal-containing salt, the third metal containing salt, . . . the N-containing salts are mixed into the liquid mixture. Exemplary metal-dopant-containing salts include, but not limited to, of magnesium nitrate ($\text{Mg}(\text{NO}_3)_2$), magnesium acetate (MgAc , $\text{Mg}(\text{CH}_3\text{COO})_2$), magnesium chloride (MgCl_2), magnesium sulfate (MgSO_4), magnesium formate ($\text{C}_2\text{H}_2\text{MgO}_4$), aluminum nitrate ($\text{Al}(\text{NO}_3)_3$), aluminum acetate (AlAc , $\text{C}_6\text{H}_9\text{AlO}_6$), aluminum chloride (AlCl_3), aluminum sulfate ($\text{Al}_2(\text{SO}_4)_3$), aluminum formate ($\text{Al}(\text{HCOO})_3$), manganese sulfate (MnSO_4), manganese nitrate ($\text{Mn}(\text{NO}_3)_2$), manganese acetate ($\text{Mn}(\text{CH}_3\text{COO})_2$), manganese formate ($\text{Mn}(\text{CHO}_2)_2$), manganese chloride (MnCl_2), zirconium nitrate ($\text{Zr}(\text{NO}_3)_4$), zirconium acetate ($\text{C}_8\text{H}_{12}\text{O}_8\text{Zr}$), zirconium chloride (ZrCl_4), zirconium sulfate ($\text{Zr}(\text{SO}_4)_2$), zirconium formate ($\text{C}_4\text{H}_4\text{O}_8\text{Zr}$), nickel sulfate (NiSO_4), nickel nitrate ($\text{Ni}(\text{NO}_3)_2$), nickel acetate ($\text{Ni}(\text{CH}_3\text{COO})_2$), nickel formate ($\text{Ni}(\text{CHO}_2)_2$), nickel chloride (NiCl_2), titanyl nitrate ($\text{TiO}(\text{NO}_3)_2$), aluminum (Al)-containing compound, magnesium (Mg)-containing compound, titanium (Ti)-containing compound, sodium (Na)-containing compound, potassium (K)-containing compound, scandium (Sc)-containing compound, niobium (Nb)-containing compound, neodymium (Nd)-containing compound, lanthanum (La)-containing compound, cerium (Ce)-containing compound, silicon (Si)-containing compound, rubidium (Rb)-containing compound, vanadium (V)-containing compound, cesium (Cs)-containing compound, chromium (Cr)-containing compound, copper (Cu)-containing compound, magnesium (Mg)-containing compound, manganese (Mn)-containing compound, zirconium (Zr)-containing compound, zinc (Zn)-containing compound, tin (Sn)-containing compound, gallium (Ga)-containing compound, barium (Ba)-containing compound, actinium (Ac)-containing compound, calcium (Ca)-containing compound, iron (Fe)-containing compound, boron (B)-containing compound, germanium (Ge)-containing compound, arsenic (As)-

containing compound, hafnium (Hf)-containing compound, Molybdenum (Mo)-containing compound, tungsten (W)-containing compound, rhenium (Re)-containing compound, ruthenium (Ru)-containing compound, rhodium (Rh)-containing compound, platinum (Pt)-containing compound, silver (Ag)-containing compound, osmium (Os)-containing compound, iridium (Ir)-containing compound, gold (Au)-containing compound, and combinations thereof, among others. [0067] Not wishing to be bound by theory, it is contemplated that, all the required metal-containing salts are first prepared in liquid phase (e.g., into a solution, slurry, or gel-like mixtures) using the lithium-containing salt, the cobalt-containing salt, the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . the N-containing salt as the sources of each metal element such that the different metals can be mixed uniformly at desired ratio. As an example, to prepare a liquid mixture of an aqueous solution, slurry or gel, one or more metal dopants with high water solubility can be used. For example, metal nitrate, metal sulfate, metal chloride, metal acetate, and metal format, etc., can be used. Organic solvents, such as alcohols, isopropanol, etc., can be used to dissolve and/or disperse metal-containing salt compounds with low water solubility. In some cases, the pH value of the liquid mixture can be adjusted to increase the solubility of the one or more precursor compounds. Optionally, chemical additives, gelation agents, and surfactants, such as ammonia, EDTA, etc., can be added into the liquid mixture to help dissolve or disperse the compounds in a chosen solvent.

[0068] At step **130**, the mist of the liquid mixture is mixed with a gas flow of a gas inside a mist generator to form a gas-liquid mixture. In addition, the liquid mixture is mixed with a gas flow of another gas inside a drying chamber. It is contemplated that these gas flows are provided to thoroughly mix the liquid mixture to uniformly form into the gas-liquid mixture and assist in carrying the gas-liquid mixture inside the drying chamber. The method **100** further includes a step **140** of drying the gas-liquid mixture at a drying temperature in the presence of the gas flows for a time period to obtain gas-solid mixtures.

[0069] The gases within the gas flows may be, for example, air, oxygen, carbon dioxide, nitrogen gas, hydrogen gas, inert gas, noble gas, and combinations thereof, among others. The gas flows may be pumped through an air filter to remove any particles, droplets, or contaminants, and the flow rate of the gases can be adjusted by a valve or other means. Accordingly, one embodiment of the invention provides that the gases are used as the gas source for carrying out drying reaction, evaporation, dehydration, and/or other reactions. In another embodiment, the gases are heated to a drying temperature to mix with the mist and remove moisture from the mist.

[0070] The drying temperature can be, for example, about 200° C. or higher, such as from 200° C. to 300° C., or at 250° C. The time period is around 1 second to 1 hour. Optionally, additional gas flow may be used to perform the drying reaction. The additional gas may be, for example, air, oxygen, carbon dioxide, nitrogen gas, hydrogen gas, inert gas, noble gas, and combinations thereof, among others. The additional gas flow may be pumped through an air filter to remove any particles, droplets, or contaminants, and the flow rate of the additional gas can be adjusted by a valve or other means.

[0071] Next, at step **150**, step **150** includes separating the gas-solid mixture into one or more solid particles of an oxide material and waste products. The gas-solid mixture comprising of the gas and the compounds mixed together are separated into one or more solid particles of oxide materials and waste products. The one or more solid particles of the oxide material may include thoroughly mixed solid particles of the compounds. Accordingly, the step **150** of the method **100** of preparing a battery material includes obtaining one or more solid particles of the oxide material from a gas-solid mixture comprised of a gas and one or more compounds.

[0072] The method **100** further includes a step **160** of annealing the one or more solid particles of an oxide material at an annealing temperature for a time period to obtain crystalized lithium cobalt oxide materials doped with one or more metal dopants of desired size, morphology and crystal structure with a formula of $\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}$. Math.doped $\text{Me1}_{\text{sub.a}}\text{Me2}_{\text{sub.b}}\text{Me3}_{\text{sub.c}}$.

. . MeN.sub.n, wherein the atomic ratio of Li:Co:Me1:Me2:Me3 . . . :MeN equaling to x:y:a:b:c: . . :n. The annealing temperature is from 400° C. to 1200° C., for example, more than 900° C., such as 1050° C. The time period is about 1 second to 10 hours.

[0073] FIG. 1B illustrates another embodiment of a flow chart of a method **200** of producing a lithium cobalt oxide material doped with one or more metal dopants for lithium ion batteries. The method **200** comprises a first step **210** of forming a mist of a liquid mixture having a lithium-containing salt compound, a cobalt-containing salt compound, and one or more metal-dopant-containing salts compounds at a molar ratio of

M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . :M.sub.MeNSalt, where each of the one or more metal-dopant-containing salts compounds is selected from a group consisting of a first metal-containing salt compound, a second metal-containing salt compound, a third metal-containing salt compound, . . . an N metal-containing salt compound and combinations thereof. The molar ratio

M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . :M.sub.MeNSalt is adjusted to be a ratio of about x:y:a:b:c: . . . :n for making the lithium cobalt oxide doped with one or more metal dopants (Li.sub.xCo.sub.yO.sub.z.Math.doped Me1.sub.a Me2.sub.b Me3.sub.c . . . MeN.sub.n) at desirable atomic ratio of Li:Co:Me1:Me2:Me3 . . . :MeN equaling to x:y:a:b:c: . . . :n, where x is from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and where $N \geq 1$, a is from 0 to 0.05 ($0 \leq a \leq 0.05$), b is from 0 to 0.05 ($0 \leq b \leq 0.05$), c is from 0 to 0.05 ($0 \leq c \leq 0.05$), . . . , and n is from 0 to 0.05 ($0 \leq n \leq 0.05$).

[0074] In one embodiment, the desired molar ratio of

M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . :M.sub.MeNSalt can be achieved by measuring and preparing appropriate amounts a lithium-containing salt (LiSalt), a cobalt-containing salt (CoSalt), a first metal dopant-containing salt (Me1Salt), a second metal dopant-containing salt (Me2Salt), a third metal dopant-containing salt (Me3Salt), . . . , and a N metal dopant-containing salt (MeNSalt). For example, the molar ratio

M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . :M.sub.MeNSalt of the lithium-containing salt, the cobalt-containing salt, the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . the N metal-containing salt can be adjusted (e.g., manually or digitally using a processing system of the invention) and prepared directly into a liquid mixture in a desired concentration prior to forming the mist of the liquid mixture. As another example, the adjusting the molar ratio

M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . :M.sub.MeNSalt of the lithium-containing salt, the cobalt-containing salt, the first metal-containing salt, the second metal-containing salt, the third metal-containing salt, . . . and the N-containing salt can be performed simultaneously with forming the mist of the liquid mixture.

[0075] In one embodiment, liquid form of lithium-containing salt compound, cobalt-containing salt compound and one or more metal-dopant-containing salts can be adjusted and prepared directly into a liquid mixture in a desired concentration. The liquid form of the lithium-containing salt compound, the cobalt-containing salt compound and the one or more metal-dopant-containing salts can be dissolved or dispersed in a suitable solvent (e.g., water, alcohol, methanol, isopropyl alcohol, organic solvents, inorganic solvents, organic acids, sulfuric acid (H.sub.2SO.sub.4), citric acid (C.sub.6H.sub.8O.sub.7), acetic acids (CH.sub.3COOH), butyric acid (C.sub.4H.sub.8O.sub.2), lactic acid (C.sub.3H.sub.6O.sub.3), Nitric acid (HNO.sub.3), hydrochloric acid (HCl), ethanol, pyridine, ammonia, acetone, and their combinations) to form into a liquid mixture of an aqueous solution, slurry, gel, aerosol or any other suitable liquid forms.

[0076] In another embodiment, the lithium-containing salt, the cobalt-containing salt and the one or more metal-dopant-containing salts can be used, depending on the desired composition of final solid product particles. For example, one or more solid particles of an oxide material can be digitally adjusted and prepared in desirable molar ratio and mixed into a liquid mixture, such as by

digitally adjusting, measuring and preparing appropriate amounts of the lithium-containing salt, the cobalt-containing salt and the one or more metal-dopant-containing salts into a container with suitable amounts of a solvent. Depending on the solubility of the lithium-containing salt, the cobalt-containing salt and the one or more metal-dopant-containing salts in a chosen solvent, pH, temperature, and mechanical stirring and mixing can be adjusted to obtain a liquid mixture where the one or more metal-dopant-containing salts at the desirable molar concentrations are fully dissolved and/or evenly dispersed.

[0077] In yet another embodiment, the lithium-containing salt, the cobalt-containing salt and the one or more metal-dopant-containing salts are mixed into a liquid mixture for obtaining final solid product particles of a mixed metal oxide material.

[0078] For example, the lithium containing salts and the cobalt containing salts are mixed into the liquid mixture. Exemplary lithium containing salts include, but not limited to, lithium sulfate (Li.sub.2SO.sub.4), lithium nitrate (LiNO.sub.3), lithium carbonate (Li.sub.2CO.sub.3), lithium acetate (LiCH.sub.2COO), lithium hydroxide (LiOH), lithium formate (LiCHO.sub.2), lithium chloride (LiCl), and combinations thereof. Exemplary cobalt containing salts include, but not limited to, cobalt sulfate (CoSO.sub.4), cobalt nitrate (Co(NO.sub.3).sub.2), cobalt acetate (Co(CH.sub.2COO).sub.2), cobalt formate (Co(CHO.sub.2).sub.2), cobalt chloride (CoCl.sub.2), and combinations thereof.

[0079] As another example, the one or more metal-dopant-containing salts are mixed into the liquid mixture. Exemplary other metal-containing salts include, but not limited to, of magnesium nitrate Mg(NO.sub.3).sub.2, magnesium acetate (MgAc, Mg(CH.sub.3COO).sub.2), magnesium chloride (MgCl.sub.2), magnesium sulfate (MgSO.sub.4), magnesium formate (C.sub.2H.sub.2MgO.sub.4), aluminum nitrate (Al(NO.sub.3).sub.3), aluminum acetate (AlAc, C.sub.6H.sub.9AlO.sub.6), aluminum chloride (AlCl.sub.3), aluminum sulfate (Al.sub.2(SO.sub.4).sub.3), aluminum formate (Al(HCOO).sub.3), manganese sulfate (MnSO.sub.4), manganese nitrate (Mn(NO.sub.3).sub.2), manganese acetate (Mn(CH.sub.2COO).sub.2), manganese formate (Mn(CHO.sub.2).sub.2), manganese chloride (MnCl.sub.2), zirconium nitrate (Zr(NO.sub.3).sub.4), zirconium acetate (C.sub.8H.sub.12O.sub.8Zr), zirconium chloride (ZrCl.sub.4), zirconium sulfate (Zr(SO.sub.4).sub.2), zirconium formate (C.sub.4H.sub.4O.sub.8Zr), nickel sulfate (NiSO.sub.4), nickel nitrate (Ni(NO.sub.3).sub.2), nickel acetate (Ni(CH.sub.2COO).sub.2), nickel formate (Ni(CHO.sub.2).sub.2), nickel chloride (NiCl.sub.2), titanyl nitrate ((TiO(NO.sub.3).sub.2)), aluminum (Al)-containing compound, magnesium (Mg)-containing compound, titanium (Ti)-containing compound, sodium (Na)-containing compound, potassium (K)-containing compound, scandium (Sc)-containing compound, niobium (Nb)-containing compound, neodymium (Nd)-containing compound, lanthanum (La)-containing compound, cerium (Ce)-containing compound, silicon (Si)-containing compound, rubidium (Rb)-containing compound, vanadium (V)-containing compound, cesium (Cs)-containing compound, chromium (Cr)-containing compound, copper (Cu)-containing compound, magnesium (Mg)-containing compound, manganese (Mn)-containing compound, zirconium (Zr)-containing compound, zinc (Zn)-containing compound, tin (Sn)-containing compound, gallium (Ga)-containing compound, barium (Ba)-containing compound, actinium (Ac)-containing compound, calcium (Ca)-containing compound, iron (Fe)-containing compound, boron (B)-containing compound, germanium (Ge)-containing compound, arsenic (As)-containing compound, hafnium (Hf)-containing compound, Molybdenum (Mo)-containing compound, tungsten (W)-containing compound, rhenium (Re)-containing compound, ruthenium (Ru)-containing compound, rhodium (Rh)-containing compound, platinum (Pt)-containing compound, silver (Ag)-containing compound, osmium (Os)-containing compound, iridium (Ir)-containing compound, gold (Au)-containing compound, and combinations thereof, among others.

[0080] Not wishing to be bound by theory, it is contemplated that, all of the required metal elements are first mixed in liquid phase (e.g., into a solution, slurry, or gel) using metal-containing salts as the sources of each metal element such that the different metals can be mixed uniformly at

desired ratio. As an example, to prepare a liquid mixture of an aqueous solution, slurry or gel, one or more metal dopants with high water solubility can be used. For example, metal nitrate, metal sulfate, metal chloride, metal acetate, and metal format, etc., can be used. Organic solvents, such as alcohols, isopropanol, etc., can be used to dissolve and/or disperse metal-containing salt with low water solubility. In some cases, the pH value of the liquid mixture can be adjusted to increase the solubility of the one or more precursor compounds. Optionally, chemical additives, gelation agents, and surfactants, such as ammonia, EDTA, etc., can be added into the liquid mixture to help dissolve or disperse the compounds in a chosen solvent.

[0081] Secondly, at step **220** of the method **200**, the method includes flowing a flow of a gas into a drying chamber. The flow of the gas may be pumped through an air filter to remove any particles, droplets, or contaminants, and the flow rate of the gas can be adjusted by a valve or other means. In one embodiment, the gas is heated to a drying temperature to mix with the mist and remove moisture from the mist.

[0082] The mist of the liquid mixture may be generated by a mist generator, such as a nozzle, a sprayer, an atomizer, or any other mist generators. Most mist generators employ air pressure or other means to convert a liquid mixture into liquid droplets. The mist generator can be coupled to a portion of the drying chamber to generate a mist (e.g., a large collection of small size droplets) of the liquid mixture directly within the drying chamber. As an example, an atomizer can be attached to a portion of the drying chamber to spray or inject the liquid mixture into a mist containing small sized droplets directly inside the drying chamber. In general, a mist generator that generates a mist of mono-sized droplets are desirable. Alternatively, a mist can be generated outside the drying chamber and delivered into the drying chamber.

[0083] Desired liquid droplet sizes can be adjusted by adjusting the sizes of liquid delivery/injection channels within the mist generator. Droplet size ranging from a few nanometers to a few hundreds of micrometers can be generated. Suitable droplet sizes can be adjusted according to the choice of the mist generator used, the precursor compounds, the temperature of the drying chamber, the flow rate of the gas, and the residence time inside the drying chamber. As an example, a mist with liquid droplet sizes between one tenth of a micron and one millimeter is generated inside the drying chamber.

[0084] Then, at step **230** of the method **200**, a mist of the liquid mixture is mixed with the flow of a gas to form a gas-liquid mixture prior to and/or after the liquid mixture is inside the drying chamber. The mist is formed from a liquid mixture dissolved and/or dispersed in a suitable liquid solvent. The flow of one or more gases and the flow of the mist are mixed together to form a gas-liquid mixture. The gases may be, for example, air, oxygen, carbon dioxide, nitrogen gas, hydrogen gas, inert gas, noble gas, and combinations thereof, among others. The gases may be pumped through an air filter to remove any particles, droplets, or contaminants, and the flow rate of the gases can be adjusted by a valve or other means.

[0085] In one example, the mist of the liquid mixture is mixed with a flow of a carrying gas inside the mist generator prior to delivering into the drying chamber. In another example, the mist of the liquid mixture is mixed with a flow of a drying gas inside the drying chamber and carrying through the drying chamber to be dried. Accordingly, one embodiment of the invention provides that one or more gases flown within the drying chamber are used as the gas source for carrying out drying reaction, evaporation, dehydration, and/or other reactions inside the drying chamber such that gas-liquid mixtures are dried into gas-solid mixtures. In another embodiment, the gases is heated to a drying temperature to mix with the mist and remove moisture from the mist.

[0086] At step **240**, drying the gas-liquid mixture at a drying temperature in the presence of the gas and forming a gas-solid mixture is performed. The mist of the liquid mixture is dried (e.g., removing its moisture, liquid, etc.) at a drying temperature for a desired residence time and form into a gas-solid mixture with the flow of the gases within the drying chamber. As the removal of the moisture from the mist of the liquid mixture is performed within the drying chamber filled with the

gases, a gas-solid mixture comprising of the gases and the compounds is formed. Accordingly, one embodiment of the invention provides that the gases flown within the drying chamber are used as the gas source for forming a gas-solid mixture within the drying chamber. To illustrate, the liquid mixture is dried inside the drying chamber and the drying temperature inside the drying chamber is maintained via a heating element coupled to the drying chamber, where the heating element can be a suitable heating mechanism, such as wall-heated furnace, electricity powered heater, fuel-burning heater, etc.

[0087] In another embodiment, the gases flown within the drying chamber is heated and the thermal energy of the heated gas is served as the energy source for carrying out drying reaction, evaporation, dehydration, and/or other reactions inside the drying chamber. The gas can be heated to a drying temperature by passing through a suitable heating mechanism, such as electricity powered heater, fuel-burning heater, etc. The drying temperature is about 200° C. or higher, for example, from 200° C. to 300° C., such as 250° C. For instance, the liquid mixture is dried in the presence of the gas that is heated to 200° C. or higher inside the drying chamber and the gas is delivered into the drying chamber to maintain the drying temperature inside the drying chamber.

[0088] In one configuration, the gas is pre-heated to a drying temperature of about 200° C. or higher prior to flowing into the drying chamber. In another configuration, drying the mist can be carried out by heating the drying chamber directly, such as heating the chamber body of the drying chamber. For example, the drying chamber can be a wall-heated furnace to maintain the drying temperature within internal plenum of the drying chamber. The advantages of using heated gas are fast heat transfer, high temperature uniformity, and easy to scale up, among others. The drying chambers may be any chambers, furnaces with enclosed chamber body, such as a dome type ceramic drying chamber, a quartz chamber, a tube chamber, etc. Optionally, the chamber body is made of thermal insulation materials (e.g., ceramics, etc.) to prevent heat loss during drying.

[0089] The gases may be, for example, air, oxygen, carbon dioxide, nitrogen gas, hydrogen gas, inert gas, noble gas, and combinations thereof, among others. For example, heated air can be used as an inexpensive gas source and energy source for drying the mist. The choice of the gases may be a gas that mix well with the mist of the liquid mixture and dry the mist without reacting to the compounds. In some cases, the chemicals in the droplets/mist may react to the gases and/or to each other to certain extent during drying, depending on the drying temperature and the chemical composition of the compounds. In addition, the residence time of the mist of thoroughly mixed compounds within the drying chamber is adjustable and may be, for example, between one second and one hour, depending on the flow rate of the gases, and the length and volume of the path that the mist has to flow through within the drying chamber.

[0090] The gas-liquid mixture is being dried within the drying chamber using the heated gases flow continuously and/or at adjustable, variable flow rates. At the same time, dried solid particles of compounds are carried by the gases, as a thoroughly-mixed gas-solid mixture, through a path within the drying chamber, and as more gases is flown in, the gas-solid mixture is delivered out of the drying chamber and continuously delivered to a gas-solid separator connected to the drying chamber.

[0091] Not wishing to be bound by theory, in the method **200** of manufacturing a battery material using the lithium-containing salt, the cobalt-containing salt and the one or more metal-dopant-containing salts, it is contemplated that the lithium-containing salt, the cobalt-containing salt and the one or more metal-dopant-containing salts are prepared into a liquid mixture and then converted into droplets, each droplet will have the one or more liquid mixture uniformly distributed. Then, the moisture of the liquid mixture is removed by passing the droplets through the drying chamber and the flow of the gas is used to carry the mist within the drying chamber for a suitable residence time. It is further contemplated that the concentrations of the compounds in a liquid mixture and the droplet sizes of the mist of the liquid mixture can be adjusted to control the chemical composition, particle sizes, and size distribution of final solid product particles of the battery material. It is

designed to obtain spherical solid particles from a thoroughly mixed liquid mixture of two or more precursors after drying the mist of the liquid mixture. In contrast, conventional solid-state manufacturing processes involve mixing or milling a solid mixture of precursor compounds, resulting in uneven mixing of precursors.

[0092] Next, at step **250**, step **250** includes separating the gas-solid mixture into one or more solid particles of an oxide material and waste products by a gas-solid separator. The gas-solid mixture comprising of the gas and the compounds mixed together are separated into one or more solid particles of an oxide material and a waste product. The one or more solid particles of an oxide material may include thoroughly mixed solid particles of the compounds. Accordingly, the step **250** of the method **200** of preparing a battery material includes obtaining one or more solid particles of an oxide material from a gas-solid mixture comprised of a gas and one or more compounds.

[0093] In the method **200** of preparing final solid product particles of the battery material in multiple stages, it is contemplated to perform one or more reactions of the compounds in a drying stage, two or more reaction stages, one or more cooling stages, etc., in order to obtain final solid product particles of the crystalized lithium cobalt oxide materials at desired size, morphology and crystal structure, which are ready for further battery applications. Not wishing to be bound by theory, it is designed to perform the reaction of the compounds in two or more reaction stages to allow sufficient time and contact of the compounds to each other, encourage nucleation of proper crystal structure and proper folding of particle morphology, incur lower-thermodynamic energy partial reaction pathways, ensure thorough reactions of all compounds, and finalize complete reactions, among others.

[0094] The one or more solid particles of a lithium cobalt oxide material comprising the compounds are then processed in two or more processing stages using at least a reaction module designed for initiating reactions, and one or more reaction modules designed for completing reactions and obtaining final solid product particles of the crystalized lithium cobalt oxide materials. Additional reaction modules can also be used. In one embodiment, the reaction module includes one anneal reaction to react and oxidize the one or more solid particles of a lithium cobalt oxide material into an oxidized reaction product, where a portion of them are partially reacted (some complete reactions may occur). The another reaction module includes annealing the oxidized reaction product into final solid product particles of the crystalized lithium cobalt oxide materials to ensure complete reactions of all the reaction products.

[0095] Accordingly, the method **200** may include a processing stage of drying a mist of a liquid mixture and obtaining one or more solid particles of an oxide material using a processing module comprised of a drying chamber and a gas-solid separator. The method **200** may further include another processing stage of reacting, oxidizing and annealing the one or more solid particles of an oxide material using a reaction module comprised of an annealing chamber.

[0096] At step **260**, step **260** includes delivering the solid particles of the oxide material into an annealing chamber to react and anneal the solid particles of the oxide material in the presence of a flow of a gas at an annealing temperature to obtain crystalized lithium cobalt oxide materials doped with one or more metal dopants.

[0097] The one or more solid particles of an oxide material is delivered into an annealing chamber once the one or more solid particles of an oxide material are separated from the waste product. The one or more solid particles of the oxide material is reacted and oxidized in the presence of a gas within the annealing chamber to form an oxidized reaction product. Reactions of the one or more solid particles of the oxide material within the annealing chamber may include any of oxidation, reduction, decomposition, combination reaction, phase-transformation, re-crystallization, single displacement reaction, double displacement reaction, combustion, isomerization, and combinations thereof. For example, the one or more solid particles of the oxide material may be oxidized, such as oxidizing the precursor compounds into an oxide material.

[0098] Exemplary gases include, but not limited to air, oxygen, carbon dioxide, an oxidizing gas,

nitrogen gas, inert gas, noble gas, and combinations thereof. For an oxidation reaction inside the annealing chamber, such as forming an oxide material from one or more precursors, an oxidizing gas can be used as the gas for annealing. Accordingly, one embodiment of the invention provides that the gas flows within the annealing chamber is used to oxidize the one or more solid particles of the oxide material. The gases, for example, can be air or oxygen and combination thereof. If desired, the gases can be oxygen with high purity; the purity of the oxygen is more than 50%, for example more than 80%, such as 95%. Accordingly, the gas flows within the annealing chamber is served as the energy source for carrying out reaction, oxidation, and/or other reactions inside the annealing chamber.

[0099] At this stage of the process, the step **260** further includes delivering the solid particles of the oxide material into an annealing chamber to react and annealing the solid particles of the lithium cobalt oxide material in the presence of a flow of a gas at an annealing temperature of 400° C. or higher for a residence time to obtain crystalized lithium cobalt oxide materials doped with one or more metal dopants. For example, the annealing temperature can be more than 900° C., such as 1050° C., such as 1000° C. The residence time is about 1 second to 10 hours.

[0100] In one embodiment, the gas flown within the annealing chamber is heated and the thermal energy of the heated gas is served as the energy source for carrying out annealing reaction, and/or other reactions inside the annealing chamber. The gas can be heated to a temperature of 550° C. or higher by passing through a suitable heating mechanism, such as electricity powered heater, fuel-burning heater, etc. For instance, the one or more solid particles of the oxide materials are annealed in the presence of the gas that is heated to 550° C. or higher inside the annealing chamber and the gas is delivered into the annealing chamber to maintain the annealing temperature inside the annealing chamber.

[0101] Another embodiment of the present invention is that the one or more solid particles of the oxide materials are annealed inside the annealing chamber and the annealing temperature inside the annealing chamber is maintained via a heating element coupled to the annealing chamber, where the heating element can be a suitable heating mechanism, such as wall-heated furnace, electricity powered heater, fuel-burning heater, etc.

[0102] In one configuration, the gas is pre-heated to a temperature of about 550° C. or higher prior to flowing into the annealing chamber. In another configuration, annealing the one or more solid particles of the oxide materials can be carried out by heating the annealing chamber directly, such as heating the chamber body of the annealing chamber. For example, the annealing chamber can be a wall-heated furnace to maintain the annealing temperature within internal plenum of the annealing chamber. The advantages of using heated gas are fast heat transfer, high temperature uniformity, and easy to scale up, among others. The annealing chambers may be any chambers, furnaces with enclosed chamber body, such as a dome type ceramic annealing chamber, a quartz chamber, a tube chamber, etc. Optionally, the chamber body of the annealing chamber is made of thermal insulation materials (e.g., ceramics, etc.) to prevent heat loss during annealing process.

[0103] The gas may be, for example, air, oxygen, carbon dioxide, nitrogen gas, hydrogen gas, inert gas, noble gas, and combinations thereof, among others. For example, heated air can be used as an inexpensive gas source and energy source for drying the mist. In addition, the residence time within the annealing chamber is adjustable and may be, for example, between one second and one hour, depending on the flow rate of the gas, and the length and volume of the path that the solid particles have to pass through within the annealing chamber.

[0104] The method **200** may include a processing stage of cooling the crystalized lithium cobalt oxide materials doped with one or more metal dopants and obtaining final solid product particles of the crystalized lithium cobalt oxide materials doped with one or more metal dopants at desired size, morphology and crystal structure at step **270**. For example, the temperature of the final solid product particles of the crystalized lithium cobalt oxide materials doped with one or more metal dopants may be slowly cooled down to room temperature to avoid interfering or destroying a

process of forming into its stable energy state with uniform morphology and desired crystal structure. In another example, the cooling stage may be performed very quickly to quench the reaction product such the crystal structure of the solid particles of the reaction product can be formed at its stable energy state. As another example, a cooling processing stage in a multi-stage continuous process may include a cooling module comprised of one or more cooling mechanisms. Exemplary cooling mechanisms may be, for example, a gas-solid separator, a heat exchanger, a gas-solid feeder, a fluidized bed cooling mechanism, and combinations thereof, among others.

[0105] FIG. 2 illustrates a flow chart of incorporating the method **100** of preparing a material for a battery electrochemical cell using a system **300** fully equipped with all of the required manufacturing tools. The system **300** generally includes a mist generator **306**, a drying chamber **310**, a gas-solid separator **320**, and a reactor **340**. First, a liquid mixture containing two or more precursors is prepared and delivered into the mist generator **306** of the system **300**. The mist generator **306** is coupled to the drying chamber **310** and adapted to generate a mist from the liquid mixture. A flow of heated gases can be flowed into the drying chamber **310** to fill and pre-heat an internal volume of the drying chamber **310** prior to the formation of the mist or at the same time when the mist is generated inside the drying chamber **310**. The mist is mixed with the heated gas and its moisture is removed such that a gas-solid mixture, which contains the heated gases, two or more precursors, and/or other gas-phase waste product or by-products, etc., is formed.

[0106] Next, the gas-solid mixture is continuously delivered into the gas-solid separator **320** which separates the gas-solid mixture into solid particles and waste products. The solid particles is then delivered into the reactor **340** to be mixed with a flow of heated gas and form a gas-solid mixture. The reaction inside the reactor is carried out for a reaction time until reaction products can be obtained. Optionally, the reaction product gas-solid mixture can be delivered into a gas-solid separator (e.g., a gas-solid separator **328**) to separate and obtain final solid product particles and a gaseous side product. In addition, one or more flows of cooling fluids (e.g., gases or liquids) may be used to cool the temperature of the reaction products. The final solid product particles can be delivered out of the system **300** for further analysis on their properties (e.g., specific capacity, power performance, battery charging cycle performance, etc.), particle sizes, morphology, crystal structure, etc., to be used as a material in a battery cell. Finally, the final particles are packed into a component of a battery cell.

[0107] FIG. 3 is a schematic of the system **300**, which is one example of an integrated tool/apparatus that can be used to carry out a fast, simple, continuous and low cost manufacturing process for preparing a material for a battery electrochemical cell. The system **300** is connected to a liquid mixer **304**, which in turn is connected to two or more reactant sources **302A**, **302B**. The reactant sources **302A**, **302B** are provided to store various precursor compounds and liquid solvents. Desired amounts of precursor compounds (in solid or liquid form) and solvents are dosed and delivered from the reactant sources **302A**, **302B** to the liquid mixer **304** so that the precursor compounds can be dissolved and/or dispersed in the solvent and mix well into a liquid mixture. If necessary, the liquid mixer **304** is heated to a temperature, such as between 30° C. and 90° C. to help uniformly dissolve, disperse, and/or mix the precursors. The liquid mixer **304** is optionally connected to a pump **305**, which pumps the liquid mixture from the liquid mixer **304** into the mist generator **306** of the system **300** to generate a mist.

[0108] The mist generator **306** converts the liquid mixture into a mist with desired droplet size and size distribution. In addition, the mist generator **306** is coupled to the drying chamber **310** in order to dry and remove moisture from the mist and obtain thoroughly-mixed solid precursor particles. In one embodiment, the mist generator **306** is positioned near the top of the drying chamber **310** that is positioned vertically (e.g., a dome-type drying chamber, etc.) to inject the mist into the drying chamber **310** and pass through the drying chamber vertically downward. Alternatively, the mist generator can be positioned near the bottom of the drying chamber **310** that is vertically positioned to inject the mist upward into the drying chamber to increase the residence time of the mist

generated therein. In another embodiment, when the drying chamber **310** is positioned horizontally (e.g., a tube drying chamber, etc.) and the mist generator **306** is positioned near one end of the drying chamber **310** such that a flow of the mist, being delivered from the one end through another end of the drying chamber **310**, can pass through a path within the drying chamber **310** for the length of its residence time.

[0109] The drying chamber **310** generally includes a chamber inlet **315**, a chamber body **312**, and a chamber outlet **317**. In one configuration, the mist generator **306** is positioned inside the drying chamber **310** near the chamber inlet **315** and connected to a liquid line **303** adapted to flow the liquid mixture therein from the liquid mixer **304**. For example, the liquid mixture within the liquid mixer **304** can be pumped by the pump **305** through the liquid line **303** connected to the chamber inlet **315** into the internal volume of the drying chamber **310**. Pumping of the liquid mixture by the pump **305** can be configured, for example, continuously at a desired delivery rate (e.g., adjusted by a metered valve or other means) to achieve good process throughput of system **300**. In another configuration, the mist generator **306** is positioned outside the drying chamber **310** and the mist generated therefrom is delivered to the drying chamber **310** via the chamber inlet **315**.

[0110] One or more gas lines (e.g., gas lines **331A**, **331B**, **331C**, **331D**, etc.) can be coupled to various portions of the drying chamber **310** and adapted to flow a gas from a gas source **332** into the drying chamber **310**. A flow of the gas stored in the gas source **332** can be delivered, concurrently with the formation of the mist inside drying chamber **310**, into the drying chamber **310** to carry the mist through the drying chamber **310**, remove moisture from the mist, and form a gas-solid mixture containing the precursors. Also, the flow of the gas can be delivered into the drying chamber **310** prior to the formation of the mist to fill and preheat an internal volume of the drying chamber **310** prior to generating the mist inside the drying chamber **310**.

[0111] In one example, the gas line **331A** is connected to the top portion of the drying chamber **310** to deliver the gas into the mist generator **306** positioned near the chamber inlet **315** to be mixed with the mist generated by the mist generator **306** inside the drying chamber **310**. In one embodiment, the gas is preheated to a temperature of between 70° C. and 600° C. to mix with and remove moisture from the mist. As another example, the gas line **331B** delivering the gas therein is connected to the chamber inlet **315** of the drying chamber **310**, in close proximity with the liquid line **303** having the liquid mixture therein. Accordingly, the gas can thoroughly mix with the mist of the liquid mixture inside the drying chamber **310**.

[0112] In another example, the gas line **331C** is connected to the chamber body **312** of the drying chamber **310** to deliver the gas therein and mix the gas with the mist generated from the mist generator **306**. In addition, the gas line **331D** connected to the drying chamber **310** near the chamber outlet **317** may be used to ensure the gas-solid mixture formed within the drying chamber **310** is uniformly mixed with the gas.

[0113] The flow of the gas may be pumped through an air filter to remove any particles, droplets, or contaminants, and the flow rate of the gas can be adjusted by a valve or other means. In one embodiment, the gas is heated to a drying temperature to mix with the mist and remove moisture from the mist. It is designed to obtain spherical solid particles from a thoroughly-mixed liquid mixture of two or more precursors after drying the mist of the liquid mixture. In contrast, conventional solid-state manufacturing processes involve mixing or milling a solid mixture of precursor compounds, resulting in uneven mixing of precursors.

[0114] Once the mist of the liquid mixture is dried and formed into a gas-solid mixture with the gas, the gas-solid mixture is delivered out of the drying chamber **310** via the chamber outlet **317**. The drying chamber **310** is coupled to the gas-solid separator **320** of the system **300**. The gas-solid separator **320** collects chamber products (e.g., a gas-solid mixture having the gas and the one or more solid particles of a lithium cobalt oxide material mixed together) from the chamber outlet **317**.

[0115] The gas-solid separator **320** includes a separator inlet **321A**, two or more separator outlets

322A, 324A. The separator inlet **321A** is connected to the chamber outlet **317** and adapted to collect the gas-solid mixture and other chamber products from the drying chamber **310**. The gas-solid separator **320** separates the gas-solid mixture from the drying chamber **310** into one or more solid particles of a lithium cobalt oxide material and waste products. The separator outlet **322A** is adapted to deliver the one or more solid particles of a lithium cobalt oxide material to the reactor **340** for further processing and reactions. The separator outlet **324A** is adapted to deliver waste products out of the gas-solid separator **320**.

[0116] The waste products may be delivered into a gas abatement device **326A** to be treated and released out of the system **300**. The waste product may include, for example, water (H.sub.2O) vapor, organic solvent vapor, nitrogen-containing gas, oxygen-containing gas, O.sub.2, O.sub.3, nitrogen gas (N.sub.2), NO, NO.sub.2, NO.sub.2, N.sub.2O, N.sub.4O, NO.sub.3, N.sub.2O.sub.3, N.sub.2O.sub.4, N.sub.2O.sub.5, N(NO.sub.2).sub.3, carbon-containing gas, carbon dioxide (CO.sub.2), CO, hydrogen-containing gas, H.sub.2, chlorine-containing gas, Cl.sub.2, sulfur-containing gas, SO.sub.2, small particles of the one or more solid particles of a lithium cobalt oxide material, and combinations thereof.

[0117] The one or more solid particles of a lithium cobalt oxide material may include at least particles of the two or more precursors that are dried and uniformly mixed together. It is contemplated to separate the one or more solid particles of a lithium cobalt oxide material away from any side products, gaseous products or waste products, prior to reacting the two or more precursors in the reactor **340**. Accordingly, the system **300** is designed to mix the two or more precursors uniformly, dry the two or more precursors, separate the dried two or more precursors, and react the two or more precursors into final solid product particles of the crystalized lithium cobalt oxide materials in a continuous manner.

[0118] Suitable gas-solid separators include cyclones, electrostatic separators, electrostatic precipitators, gravity separators, inertia separators, membrane separators, fluidized beds, classifiers, electric sieves, impactors, particles collectors, leaching separators, elutriators, air classifiers, leaching classifiers, and combinations thereof, among others.

[0119] Once the one or more solid particles of a lithium cobalt oxide material are separated and obtained, it is delivered into the reactor **340** for further reaction. The reactor **340** includes a gas inlet **333**, a reactor inlet **345**, and a reactor outlet **347**. The reactor inlet **345** is connected to the separator outlet **322A** and adapted to receive the solid particles. Optionally, a vessel **325** is adapted to store the solid particles prior to adjusting the amounts of the solid particles delivered into the reactor **340**.

[0120] The gas inlet **333** of the reactor **340** is coupled to a heating mechanism **380** to heat a gas from a gas source **334** to an annealing temperature of between 400° C. and 1200° C. The heating mechanism **380** can be, for example, an electric heater, a gas-fueled heater, a burner, among other heaters. Additional gas lines can be used to deliver heated air or gas into the reactor **340**, if needed. The pre-heated gas can fill the reactor **340** and maintained the internal temperature of the reactor **340**, much better and energy efficient than conventional heating of the chamber body of a reactor.

[0121] The gas flown inside the reactor **340** is designed to be mixed with the one or more solid particles of a lithium cobalt oxide material and form an oxidized reaction product inside the reactor **340**. Thermal energy from the pre-heated gas is used as the energy source for reacting the one or more solid particles of a lithium cobalt oxide material within the reactor **340**. The reaction process includes, but not limited to, reduction, decomposition, combination reaction, phase-transformation, re-crystallization, single displacement reaction, double displacement reaction, combustion, isomerization, and combinations thereof. The oxidized reaction product is then going through annealing process for a residence time of between 1 second and ten hours, or longer, depending on the annealing temperature and the type of the precursors initially delivered into the system **300**. One embodiment of the invention provides the control of the temperature of the reactor **340** by the temperature of the heated gas. The use of the heated gas as the energy source inside the reactor **340**

provides the benefits of fast heat transfer, precise temperature control, uniform temperature distribution therein, and/or easy to scale up, among others.

[0122] Once the reactions inside the reactor **340** are complete, for example, upon the formation of desired crystal structure, particle morphology, and particle size, oxidized reaction products are delivered out of the reactor **340** via the reactor outlet **347** and/or a reactor outlet **348**. The cooled reaction products include final solid product particles of the crystalized lithium cobalt oxide materials containing, for example, oxidized reaction product particles of the precursor compounds which are suitable as a material of a battery cell.

[0123] Optionally, the system **300** includes a gas-solid separator, such as a gas-solid separator **328**, which collects the reaction products from the reactor outlet **347** of the reactor **340**. The gas-solid separator **328** may be a particle collector, such as cyclone, electrostatic separator, electrostatic precipitator, gravity separator, inertia separator, membrane separator, fluidized beds classifiers electric sieves impactor, leaching separator, elutriator, air classifier, leaching classifier, and combinations thereof.

[0124] The gas-solid separator **328** of the system **300** generally includes a separator inlet **321B**, a separator outlet **322B** and a separator outlet **324B** and is used to separate the reaction products into the solid particles and gaseous side products. The gaseous side products may be delivered into a gas abatement device **326B** to be treated and released out of the system **300**. The gaseous side products separated by the gas-solid separator **328** may generally contain water (H.sub.2O) vapor, organic solvent vapor, nitrogen-containing gas, oxygen-containing gas, O.sub.2, O.sub.3, nitrogen gas (N.sub.2), NO, NO.sub.2, NO.sub.2, N.sub.2O, N.sub.4O, NO.sub.3, N.sub.2O.sub.3, N.sub.2O.sub.4, N.sub.2O.sub.5, N(NO.sub.2).sub.3, carbon-containing gas, carbon dioxide (CO.sub.2), CO, hydrogen-containing gas, H.sub.2, chlorine-containing gas, Cl.sub.2, sulfur-containing gas, SO.sub.2, small particles of the solid particles, and combinations thereof. In addition, the system **300** may further include one or more cooling fluid lines **353**, **355** connected to the reactor outlet **347** or the separator outlet **322A** of the gas solid separator **328** and adapted to cool the reaction products and/or the solid particles. The cooling fluid line **353** is adapted to deliver a cooling fluid (e.g., a gas or liquid) from a source **352** to the separator inlet **321B** of the gas-solid separator **328**. The cooling fluid line **355** is adapted to deliver a cooling fluid, which may filtered by a filter **354** to remove particles, into a heat exchanger **350**.

[0125] The heat exchanger **350** is adapted to collect and cool the solid particles and/or reaction products from the gas-solid separator **328** and/or the reactor **340** by flowing a cooling fluid through them. The cooling fluid has a temperature lower than the temperature of the reaction products and the solid particles delivered from the gas-solid separator **328** and/or the reactor **340**. The cooling fluid may have a temperature of between 4° C. and 30° C. The cooling fluid may be liquid water, liquid nitrogen, an air, an inert gas or any other gas which would not react to the reaction products.

[0126] Final solid products particles are collected and cooled by one or more separators, cooling fluid lines, and/or heat exchangers, and once cooled, the solid particles are delivered out of the system **300** and collected in a final product collector **368**. The solid particles may include oxidized form of precursors, such as an oxide material, suitable to be packed into a battery cell **370**.

Additional pumps may also be installed to achieve the desired pressure gradient.

[0127] A process control system **390** can be coupled to the system **300** at various locations to automatically control the manufacturing process performed by the system **300** and adjust various process parameters (e.g., flow rate, mixture ratio, temperature, residence time, etc.) within the system **300**. For example, the flow rate of the liquid mixture into the system **300** can be adjusted near the reactant sources **302A**, **302B**, the liquid mixer **304**, or the pump **305**. As another example, the droplet size and generation rate of the mist generated by the mist generator **306** can be adjusted. In addition, flow rate and temperature of various gases flown within the gas lines **331A**, **331B**, **331C**, **331D**, **333**, **353**, **355**, **515**, etc., can be controlled by the process control system **390**. In addition, the process control system **390** is adapted to control the temperature and the residence

time of various gas-solid mixture and solid particles at desired level at various locations.

[0128] Accordingly, a continuous process for producing a material of a battery cell using a system having a mist generator, a drying chamber, one or more gas-solid separators and a reactor is provided. A mist generated from a liquid mixture of one or more metal precursor compounds in desired ratio is mixed with air and dried inside the drying chamber, thereby forming gas-solid mixtures. One or more gas-solid separators are used in the system to separate the gas-solid mixtures from the drying chamber into solid particles packed with the one or more metal precursors and continuously deliver the solid particles into the reactor for further reaction to obtain final solid material particles with desired ratio of two or more intercalated metals.

[0129] In one embodiment, preparation and manufacturing of a metal oxide material is provided. Depending on the details and ratios of the metal precursor compounds that are delivered into the system **300**, the resulting final solid material particles obtained from the system **300** may be a metal oxide material, a doped metal oxide material, an inorganic metal salts, among others. In addition, the metal oxide materials can exhibit a crystal structure of metals in the shape of layered, spinel, olivine, etc. In addition, the morphology of the final solid product particles (such as the oxidized reaction product prepared using the method **100** and the system **300** as described herein) exists as desired solid powders. The particle sizes of the solid powders range between 10 nm and 100 μm .

[0130] In operation, a mist is mixed with a gas flow of a gas inside a mist generator to form a gas-liquid mixture, where the liquid mixture includes a lithium-containing salt compound, a cobalt-containing salt compound, and one or more metal-dopant-containing salts compounds. In addition, the liquid mixture is mixed with a gas flow of another gas inside a drying chamber. It is contemplated that these gas flows are provided to thoroughly mix the liquid mixture to uniformly form into the gas-liquid mixture and assist in carrying the gas-liquid mixture inside the drying chamber. The liquid mixture can be adjusted digitally or manually prepared in a desirable molar ratio of the lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-dopant-containing salts compounds at a ratio of around $x:y:a:b: \dots :n$ inside reactant sources and delivered into one or more liquid mixers.

[0131] In one embodiment, the adjusting of the molar ratio of the lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-dopant-containing salt compounds is performed prior to the forming the mist of the liquid mixture inside a liquid mixer. Desired molar ratio of the lithium-containing salt, the cobalt-containing salt, and the one or more metal-dopant-containing salts are digitally or manually measured and delivered from reactant sources to the liquid mixer so that the lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-dopant-containing salts compounds can be dissolved and/or dispersed in the solvent and mix well into the liquid mixture inside the liquid mixer. The lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-containing salts compounds are all soluble in a suitable solvent within the liquid mixture.

[0132] In another embodiment, the adjusting of the molar ratio of the lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-dopant-containing salt compounds is performed simultaneously with the forming the mist of the liquid mixture. The desirable molar ratio of the lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-dopant-containing salt compounds can be adjusted digitally or manually from each reactant source and delivered into the mist generator to generate the mist of the liquid mixture inside the mist generator.

[0133] The liquid mixture comprising the lithium-containing salt compound, the cobalt-containing salt compound, and the one or more metal-containing salts compounds is mixed with a gas flow to form a gas-liquid mixture inside a drying chamber. Then, the gas-liquid mixture is dried at a drying temperature inside the drying chamber to form a gas-solid mixture of solid particles of an oxide material. The gas-solid mixture is continuously delivered into the gas-solid separator which separates the gas-solid mixture into one or more solid particles of the oxide material and waste

products.

[0134] The one or more solid particles of the oxide material are then delivered into an annealing chamber to be mixed with a flow of a gas. The one or more solid particles of the oxide material are reacted and annealed at an annealing temperature inside the annealing chamber to obtain high quality lithium cobalt oxide materials doped with one or more metal dopants at desired size, morphology and crystal structure.

Example

[0135] SUITABLE EXAMPLES: Exemplary material compositions and formulations of the present inventions are shown in Table 1. In group of A (Example #: A2-A8), lithium cobalt oxides materials doped with one or more metal dopants having a chemical formula of $\text{Li}_{x-y-z}\text{Co}_y\text{O}_z\text{Math.doped Me1}_a\text{Me2}_b\text{Me3}_c \dots \text{MeN}_n$, is designed and prepared such that a ratio of $x:y:a:b:c : \dots n$ is equivalent to $M_{\text{LiSalt}}:M_{\text{CoSalt}}:M_{\text{Me1Salt}}:M_{\text{Me2Salt}}:M_{\text{Me3Salt}}: \dots M_{\text{MeNSalt}}$, wherein x is from 0.95-0.99 ($0.95 \leq x \leq 0.99$), y is 1.0, the a is from 0-0.05 ($0 \leq a \leq 0.05$), the b is from 0-0.05 ($0 < b \leq 0.05$), the c is from 0-0.05 ($0 \leq c \leq 0.05$), $N \geq 1$, and wherein Me1 , Me2 , Me3 are different metal dopants incorporated into lithium cobalt oxide materials.

[0136] For example, each of the one or more metal dopants (i.e. Me1 , Me2 , Me3 , $\dots \text{MeN}$) can be selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

[0137] For example, in group A, exemplary lithium-containing salt compounds include, but not limited to lithium sulfate ($\text{Li}_{2.4}\text{SO}_4$), lithium nitrate (LiNO_3), lithium carbonate (Li_2CO_3), lithium acetate (LiCH_2COO), lithium hydroxide (LiOH), lithium formate (LiCHO_2), lithium chloride (LiCl), and combinations thereof. Exemplary cobalt-containing salt compounds include, but not limited to cobalt containing salts include, but not limited to, cobalt sulfate (CoSO_4), cobalt nitrate (Co(NO)_3), cobalt acetate ($\text{Co(CH}_2\text{COO)}_2$), cobalt formate (Co(CHO)_2), cobalt chloride (CoCl_2), and combinations thereof.

[0138] Exemplary metal-dopant-containing salts include, but not limited to, of magnesium nitrate (Mg(NO)_3), magnesium acetate (MgAc , $\text{Mg(CH}_2\text{COO)}_2$), magnesium chloride (MgCl_2), magnesium sulfate (MgSO_4), magnesium formate ($\text{C}_2\text{H}_2\text{MgO}_4$), aluminum nitrate (Al(NO)_3), aluminum acetate (AlAc , $\text{C}_6\text{H}_9\text{AlO}_6$), aluminum chloride (AlCl_3), aluminum sulfate ($\text{Al}_2(\text{SO}_4)_3$), aluminum formate (Al(HCOO)_3), manganese sulfate (MnSO_4), manganese nitrate (Mn(NO)_3), manganese acetate ($\text{Mn(CH}_2\text{COO)}_2$), manganese formate (Mn(CHO)_2), manganese chloride (MnCl_2), zirconium nitrate (Zr(NO)_3), zirconium acetate ($\text{C}_8\text{H}_{12}\text{O}_8\text{Zr}$), zirconium chloride (ZrCl_4), zirconium sulfate (Zr(SO)_4), zirconium formate ($\text{C}_4\text{H}_4\text{O}_8\text{Zr}$), nickel sulfate (NiSO_4), nickel nitrate (Ni(NO)_3), nickel acetate ($\text{Ni(CH}_2\text{COO)}_2$), nickel formate (Ni(CHO)_2), nickel chloride (NiCl_2), titanyl nitrate (TiO(NO)_3). The annealing temperature and annealing time in group A experiments can be controlled from 900 to 949° C. for 15 to 20 hours.

TABLE-US-00001 TABLE 1 Exemplary LCO materials doped with one or more metal dopants compositions Example Anneal Temp Anneal Time # M_{LiSalt} M_{CoSalt} M_{Me1Salt} M_{Me2Salt} M_{Me3Salt} (° C.) (hour)

A2	0.95-0.99	1.0	0-0.05	0	0	900-949	15-20
A3	0.95-0.99	1.0	0-0.05	0	0	900-949	15-20
A4	0.95-0.99	1.0	0-0.05	0	0	900-949	15-20
A5	0.95-0.99	1.0	0-0.05	0-0.05	0	900-949	15-20
A6	0.95-0.99	1.0	0-0.05	0	0-0.05	900-949	15-20
A7	0.95-0.99	1.0	0-0.05	0	0-0.05	900-949	15-20
A8	0.95-0.99	1.0	0-0.05	0-0.05	0-0.05	900-949	15-20

[0139] Additional material compositions and formulations are shown in Table 2. In group of B (Example #: B2-B8), lithium cobalt oxides materials doped with one or more metal dopants having a chemical formula of $\text{Li}_{x-y-z}\text{Co}_y\text{O}_z\text{Math.doped Me1}_a\text{Me2}_b\text{Me3}_c \dots$

MeN.sub.n, is designed and prepared such that a ratio of x:y:a:b:c: . . . n is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . M.sub.MeNSalt, wherein x is from 0.95-0.99 ($0.95 \leq x \leq 0.99$), y is 1.0, the a is from 0-0.05 ($0 \leq a \leq 0.05$), the b is from 0-0.05 ($0 \leq b \leq 0.05$), the c is from 0-0.05 ($0 \leq c \leq 0.05$), $N \geq 1$, and wherein Me1, Me2, Me3 are different metal dopants incorporated into lithium cobalt oxide materials.

[0140] For example, each of the one or more metal dopants (i.e. Me1, Me2, Me3, . . . MeN) can be selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

[0141] For example, in group B, exemplary lithium-containing salt compounds include, but not limited to lithium sulfate (Li.sub.2SO.sub.4), lithium nitrate (LiNO.sub.3), lithium carbonate (Li.sub.2CO.sub.3), lithium acetate (LiCH.sub.2COO), lithium hydroxide (LiOH), lithium formate (LiCHO.sub.2), lithium chloride (LiCl), and combinations thereof. Exemplary cobalt-containing salt compounds include, but not limited to cobalt containing salts include, but not limited to, cobalt sulfate (CoSO.sub.4), cobalt nitrate (Co(NO.sub.3).sub.2), cobalt acetate (Co(CH.sub.2COO).sub.2), cobalt formate (Co(CHO.sub.2).sub.2), cobalt chloride (CoCl.sub.2), and combinations thereof.

[0142] Exemplary metal-dopant-containing salts include, but not limited to, of magnesium nitrate Mg(NO.sub.3).sub.2, magnesium acetate (MgAc, Mg(CH.sub.3COO).sub.2), magnesium chloride (MgCl.sub.2), magnesium sulfate (MgSO.sub.4), magnesium formate (C.sub.2H.sub.2MgO.sub.4), aluminum nitrate (Al(NO.sub.3).sub.3), aluminum acetate (AlAc, C.sub.6H.sub.9AlO.sub.6), aluminum chloride (AlCl.sub.3), aluminum sulfate (Al.sub.2(SO.sub.4).sub.3), aluminum formate (Al(HCOO).sub.3), manganese sulfate (MnSO.sub.4), manganese nitrate (Mn(NO.sub.3).sub.2), manganese acetate (Mn(CH.sub.2COO).sub.2), manganese formate (Mn(CHO.sub.2).sub.2), manganese chloride (MnCl.sub.2), zirconium nitrate (Zr(NO.sub.3).sub.4), zirconium acetate (C.sub.8H.sub.12O.sub.8Zr), zirconium chloride (ZrCl.sub.4), zirconium sulfate (Zr(SO.sub.4).sub.2), zirconium formate (C.sub.4H.sub.4O.sub.8Zr), nickel sulfate (NiSO.sub.4), nickel nitrate (Ni(NO.sub.3).sub.2), nickel acetate (Ni(CH.sub.2COO).sub.2), nickel formate (Ni(CHO.sub.2).sub.2), nickel chloride (NiCl.sub.2), titanyl nitrate (TiO(NO.sub.3).sub.2). The annealing temperature and annealing time in group B experiments can be controlled from 950 to 999° C. for 15 to 20 hours.

TABLE-US-00002 TABLE 2 Exemplary LCO materials doped with one or more metal dopants compositions Example Anneal Temp Anneal Time # M.sub.LiSalt M.sub.CoSalt M.sub.Me1Salt M.sub.Me2Salt M.sub.Me3Salt (° C.) (hour)

B2	0.95-0.99	1.0	0-0.05	0	0	950-999	15-20
B3	0.95-0.99	1.0	0	0-0.05	0	950-999	15-20
B4	0.95-0.99	1.0	0	0	0-0.05	950-999	15-20
B5	0.95-0.99	1.0	0-0.05	0-0.05	0	950-999	15-20
B6	0.95-0.99	1.0	0-0.05	0	0-0.05	950-999	15-20
B7	0.95-0.99	1.0	0	0	0-0.05	950-999	15-20
B8	0.95-0.99	1.0	0-0.05	0-0.05	0-0.05	950-999	15-20

[0143] Additional material compositions and formulations are shown in Table 3. In group of C (Example #: C2-C8), lithium cobalt oxides materials doped with one or more metal dopants having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Me1.sub.a Me2.sub.b Me3.sub.c . . . MeN.sub.n, is designed and prepared such that a ratio of x:y:a:b:c: . . . n is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . M.sub.MeNSalt, wherein x is 1.0, y is 1.0, the a is from 0-0.05 ($0 \leq a \leq 0.05$), the b is from 0-0.05 ($0 \leq b \leq 0.05$), the c is from 0-0.05 ($0 \leq c \leq 0.05$), $N \geq 1$, and wherein Me1, Me2, Me3 are different metal dopants.

[0144] For example, each of the one or more metal dopants (i.e. Me1, Me2, Me3, . . . MeN) can be selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

[0145] For example, in group C, exemplary lithium-containing salt compounds include, but not limited to lithium sulfate (Li.sub.2SO.sub.4), lithium nitrate (LiNO.sub.3), lithium carbonate (Li.sub.2CO.sub.3), lithium acetate (LiCH.sub.2COO), lithium hydroxide (LiOH), lithium formate (LiCHO.sub.2), lithium chloride (LiCl), and combinations thereof. Exemplary cobalt-containing

salt compounds include, but not limited to cobalt containing salts include, but not limited to, cobalt sulfate (CoSO₄), cobalt nitrate (Co(NO₃)₂), cobalt acetate (Co(CH₃COO)₂), cobalt formate (Co(CHO₂)₂), cobalt chloride (CoCl₂), and combinations thereof.

[0146] Exemplary metal-dopant-containing salts include, but not limited to, of magnesium nitrate Mg(NO₃)₂, magnesium acetate (MgAc, Mg(CH₃COO)₂), magnesium chloride (MgCl₂), magnesium sulfate (MgSO₄), magnesium formate (C₂H₂MgO₄), aluminum nitrate (Al(NO₃)₃), aluminum acetate (AlAc, C₆H₉AlO₆), aluminum chloride (AlCl₃), aluminum sulfate (Al₂(SO₄)₃), aluminum formate (Al(HCOO)₃), manganese sulfate (MnSO₄), manganese nitrate (Mn(NO₃)₂), manganese acetate (Mn(CH₃COO)₂), manganese formate (Mn(CHO₂)₂), manganese chloride (MnCl₂), zirconium nitrate (Zr(NO₃)₄), zirconium acetate (C₈H₁₂O₈Zr), zirconium chloride (ZrCl₄), zirconium sulfate (Zr(SO₄)₂), zirconium formate (C₄H₄O₈Zr), nickel sulfate (NiSO₄), nickel nitrate (Ni(NO₃)₂), nickel acetate (Ni(CH₃COO)₂), nickel formate (Ni(CHO₂)₂), nickel chloride (NiCl₂), titanyl nitrate (TiO(NO₃)₂). The annealing temperature and annealing time in group C experiments can be controlled from 900 to 999° C. for 15 to 20 hours.

TABLE-US-00003 TABLE 3 Exemplary LCO materials doped with one or more metal dopants compositions Exam- Anneal Anneal ple Temp Time # M.sub.LiSalt M.sub.CoSalt M.sub.Me1Salt M.sub.Me2Salt M.sub.Me3Salt (° C.) (hour) C2 1.0-1.05 1.0 0-0.05 0 0 900-999 15-20 C3 1.0-1.05 1.0 0 0-0.05 0 900-999 15-20 C4 1.0-1.05 1.0 0 0 0-0.05 900-999 15-20 C5 1.0-1.05 1.0 0-0.05 0-0.05 0 900-999 15-20 C6 1.0-1.05 1.0 0-0.05 0 0-0.05 900-999 15-20 C7 1.0-1.05 1.0 0 0-0.05 0-0.05 900-999 15-20 C8 1.0-1.05 1.0 0-0.05 0-0.05 0-0.05 900-999 15-20

[0147] Additional material compositions and formulations are shown in Table 4. In group of D (Example #: D2-D8), lithium cobalt oxides materials doped with one or more metal dopants having a chemical formula of Li_xCo_yO_z.Math.doped Me1_a Me2_b Me3_c . . . MeN_n, is designed and prepared such that a ratio of x:y:a:b:c: . . . n is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . M.sub.MeNSalt, wherein x is 1.0, y is 1.0, the a is from 0-0.05 (0≤a≤0.05), the b is from 0-0.05 (0≤b≤0.05), the c is from 0-0.05 (0≤c≤0.05), N≥1, and wherein Me1, Me2, Me3 are different metals.

[0148] For example, each of the one or more metal dopants (i.e. Me1, Me2, Me3, . . . MeN) can be selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

[0149] For example, in group D, exemplary lithium-containing salt compounds include, but not limited to lithium sulfate (Li₂SO₄), lithium nitrate (LiNO₃), lithium carbonate (Li₂CO₃), lithium acetate (LiCH₃COO), lithium hydroxide (LiOH), lithium formate (LiCHO₂), lithium chloride (LiCl), and combinations thereof. Exemplary cobalt-containing salt compounds include, but not limited to cobalt containing salts include, but not limited to, cobalt sulfate (CoSO₄), cobalt nitrate (Co(NO₃)₂), cobalt acetate (Co(CH₃COO)₂), cobalt formate (Co(CHO₂)₂), cobalt chloride (CoCl₂), and combinations thereof.

[0150] Exemplary metal-dopant-containing salts include, but not limited to, of magnesium nitrate Mg(NO₃)₂, magnesium acetate (MgAc, Mg(CH₃COO)₂), magnesium chloride (MgCl₂), magnesium sulfate (MgSO₄), magnesium formate (C₂H₂MgO₄), aluminum nitrate (Al(NO₃)₃), aluminum acetate (AlAc, C₆H₉AlO₆), aluminum chloride (AlCl₃), aluminum sulfate (Al₂(SO₄)₃), aluminum formate (Al(HCOO)₃), manganese sulfate (MnSO₄), manganese nitrate (Mn(NO₃)₂), manganese acetate (Mn(CH₃COO)₂), manganese formate (Mn(CHO₂)₂), manganese chloride (MnCl₂), zirconium nitrate (Zr(NO₃)₄), zirconium acetate

(C.sub.8H.sub.12O.sub.8Zr), zirconium chloride (ZrCl.sub.4), zirconium sulfate (Zr(SO.sub.4).sub.2), zirconium formate (C.sub.4H.sub.4O.sub.8Zr), nickel sulfate (NiSO.sub.4), nickel nitrate (Ni(NO.sub.3).sub.2), nickel acetate (Ni(CH.sub.2COO).sub.2), nickel formate (Ni(CHO.sub.2).sub.2), nickel chloride (NiCl.sub.2), titanyl nitrate (TiO(NO.sub.3).sub.2). The annealing temperature and annealing time in group A experiments can be controlled from 1000 to 1049° C. for 15 to 20 hours.

TABLE-US-00004 TABLE 4 Exemplary LCO materials doped with one or more metal dopants compositions Example Anneal Temp Anneal Time # M.sub.LiSalt M.sub.CoSalt M.sub.Me1Salt M.sub.Me2Salt M.sub.Me3Salt (° C.) (hour) D2 1.0-1.05 1.0 0-0.05 0 0 1000-1049 15-20 D3 1.0-1.05 1.0 0 0-0.05 0 1000-1049 15-20 D4 1.0-1.05 1.0 0 0 0-0.05 1000-1049 15-20 D5 1.0-1.05 1.0 0-0.05 0-0.05 0 1000-1049 15-20 D6 1.0-1.05 1.0 0-0.05 0 0-0.05 1000-1049 15-20 D7 1.0-1.05 1.0 0 0-0.05 0-0.05 1000-1049 15-20 D8 1.0-1.05 1.0 0-0.05 0-0.05 0-0.05 1000-1049 15-20

[0151] Additional material compositions and formulations are shown in Table 2. In group of E (Example #: E2-E8), lithium cobalt oxides materials doped with one or more metal dopants having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Me1.sub.a Me2.sub.b Me3.sub.c . . . MeN.sub.n, is designed and prepared such that a ratio of x:y:a:b:c: . . . n is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . M.sub.MeNSalt, Wherein x is 1.0, y is 1.0, the a is from 0-0.05 (0≤a≤0.05), the b is from 0-0.05 (0≤b≤0.05), the c is from 0-0.05 (0≤c≤0.05), N≥1, and wherein Me1, Me2, Me3 are different metal dopants.

[0152] For example, each of the one or more metal dopants (i.e. Me1, Me2, Me3, . . . MeN) can be selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

[0153] For example, in group E, exemplary lithium-containing salt compounds include, but not limited to lithium sulfate (Li.sub.2SO.sub.4), lithium nitrate (LiNO.sub.3), lithium carbonate (Li.sub.2CO.sub.3), lithium acetate (LiCH.sub.2COO), lithium hydroxide (LiOH), lithium formate (LiCHO.sub.2), lithium chloride (LiCl), and combinations thereof. Exemplary cobalt-containing salt compounds include, but not limited to cobalt containing salts include, but not limited to, cobalt sulfate (CoSO.sub.4), cobalt nitrate (Co(NO.sub.3).sub.2), cobalt acetate (Co(CH.sub.2COO).sub.2), cobalt formate (Co(CHO.sub.2).sub.2), cobalt chloride (CoCl.sub.2), and combinations thereof.

[0154] Exemplary metal-dopant-containing salts include, but not limited to, of magnesium nitrate Mg(NO.sub.3).sub.2, magnesium acetate (MgAc, Mg(CH.sub.3COO).sub.2), magnesium chloride (MgCl.sub.2), magnesium sulfate (MgSO.sub.4), magnesium formate (C.sub.2H.sub.2MgO.sub.4), aluminum nitrate (Al(NO.sub.3).sub.3), aluminum acetate (AlAc, C.sub.6H.sub.9AlO.sub.6), aluminum chloride (AlCl.sub.3), aluminum sulfate (Al.sub.2(SO.sub.4).sub.3), aluminum formate (Al(HCOO).sub.3), manganese sulfate (MnSO.sub.4), manganese nitrate (Mn(NO.sub.3).sub.2), manganese acetate (Mn(CH.sub.2COO).sub.2), manganese formate (Mn(CHO.sub.2).sub.2), manganese chloride (MnCl.sub.2), zirconium nitrate (Zr(NO.sub.3).sub.4), zirconium acetate (C.sub.8H.sub.12O.sub.8Zr), zirconium chloride (ZrCl.sub.4), zirconium sulfate (Zr(SO.sub.4).sub.2), zirconium formate (C.sub.4H.sub.4O.sub.8Zr), nickel sulfate (NiSO.sub.4), nickel nitrate (Ni(NO.sub.3).sub.2), nickel acetate (Ni(CH.sub.2COO).sub.2), nickel formate (Ni(CHO.sub.2).sub.2), nickel chloride (NiCl.sub.2), titanyl nitrate (TiO(NO.sub.3).sub.2). The annealing temperature and annealing time in group E experiments can be controlled from 1050 to 1200° C. for 15 to 20 hours.

TABLE-US-00005 TABLE 5 Exemplary LCO materials doped with one or more metal dopants compositions Example Anneal Temp Anneal Time # M.sub.LiSalt M.sub.CoSalt M.sub.Me1Salt M.sub.Me2Salt M.sub.Me3Salt (° C.) (hour) E2 1.0-1.05 1.0 0-0.05 0 0 1050-1200 15-20 E3 1.0-1.05 1.0 0 0-0.05 0 1050-1200 15-20 E4 1.0-1.05 1.0 0 0 0-0.05 1050-1200 15-20 E5 1.0-1.05 1.0 0-0.05 0-0.05 0 1050-1200 15-20 E6 1.0-1.05 1.0 0-0.05 0 0-0.05 1050-1200 15-20 E7 1.0-1.05 1.0 0 0-0.05 0-0.05 1050-1200 15-20 E8 1.0-1.05 1.0 0-0.05 0-0.05 0-0.05 1050-1200 15-20

[0155] PREPARATION: Lithium cobalt oxide materials doped with one or more metal dopants were prepared in the following steps: (a) mixing 1 M solutions of forming a liquid mixture having a lithium-containing salt at a molarity of M_{LiSalt} , a cobalt-containing salt at a molarity of M_{CoSalt} , a first metal salt at a molarity of $M_{Me1Salt}$, a second metal salt at a molarity of $M_{Me2Salt}$, and a third metal salt at a molarity of $M_{Me3Salt}$ wherein the liquid mixture achieves a molar ratio of

$M_{LiSalt}:M_{CoSalt}:M_{Me1Salt}:M_{Me2Salt}:M_{Me3Salt}$; (b) generating a mist of the liquid mixture inside a mist generator of the drying chamber. The mist of the liquid mixture is mixed with a gas flow of a gas inside a mist generator to form a gas-liquid mixture. In addition, the liquid mixture is mixed with a gas flow of another gas inside a drying chamber; (c) mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture inside the drying chamber; (d) dry the gas-liquid mixture at a drying temperature for a time period and form a gas-solid mixture inside the drying chamber; (e) separate the gas-solid mixture into one or more solid particles of a an oxide material and a waste product; (f) deliver the solid particles of the lithium cobalt oxide material into an annealing chamber to react and anneal the solid particles of the lithium cobalt oxide material in the presence of a flow of a gas at an annealing temperature to obtain crystalized lithium cobalt oxide materials doped with one or more metal dopants, and anneal the crystalized lithium cobalt oxide materials doped with one or more metal dopants inside the annealing chamber for a time period to obtain crystalized lithium cobalt oxide materials; (g) cool the crystalized lithium cobalt oxide materials doped with one or more metal dopants and obtain final solid product particles of crystalized lithium cobalt oxide materials doped with one or more metal dopants at desired size, morphology and crystal structure.

[0156] In some embodiments, the compositions and formulations of the present inventions being tested are as shown in the below Table 6. The compositions of the present inventions, prepared according to Example #12 and Example #16, have a chemical formula of

$Li_{x}Co_{y}O_{z}Math_{doped}Zr_{c}$, wherein a ratio of $x:y:c$ is equivalent to $M_{LiSalt}:M_{CoSalt}:M_{ZrSalt}$, wherein x is from 0.9-1.1 ($0.9 \leq x \leq 1.1$), x is 0.97, y is 1.0, c is 0.0017.

[0157] In Example 12 and Example 16, exemplary lithium-containing salt compounds include, but not limited to, lithium nitrate ($LiNO_{3}$), exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate ($Co(NO_{3})_{2}$) and combinations thereof, exemplary zirconium-containing salt compound include, but not limited to, zirconium nitrate ($Zr(NO_{3})_{4}$) and combinations thereof. The annealing temperature and annealing time in Example 12 and Example 16 were heated to 950° C. for 17 hours. The List of chemistries used for in the present invention is displayed in Table 6.

TABLE-US-00006 TABLE 6 Exemplary compositions of measured LCO doped with zirconium material

Example	$M_{Me1Salt}$	$M_{Me2Salt}$	M_{ZrSalt}	Anneal Temp	Anneal Time	#
M_{LiSalt}	M_{CoSalt}	$(Mg(NO_{3})_{2})$	$(Al(NO_{3})_{3})$	$(Zr(NO_{3})_{4})$	(° C.)	(hour)
12	0.97	1.0	0	0	0.0017	950 17
16	1.0	1.0	0	0	0.0017	950 17

[0158] Table 7 illustrates testing results of exemplary compositions of measured LCO material doped with zirconium (Example #12). One observation is that the testing results of the ratio of the measured LCO.Math.doped Zr material compositions of Li:Co:Zr are within an expected range from the prepared molar ratio of $M_{LiSalt}:M_{CoSalt}:M_{ZrSalt}$ being prepared.

TABLE-US-00007 TABLE 7 Exemplary compositions of measured LCO doped with zirconium material

Example #	Li	Ni	Co	Al	Mg	Zr
12	1.0312	0.0022	0.9963	0.0013	0.0001	0.0019

[0159] Table 8 illustrates testing results of electric capacity and coulombic efficiency (CE) of battery cells made by lithium cobalt oxide materials doped with zirconium at different cutoff voltage of 4.45 voltage and 4.5 voltage, prepared according to Example #16. One observation can be found that the cutoff voltage affect the initial charge and discharge capacity and CE of exemplary lithium cobalt oxide material doped with zirconium. Further observation can be found

that with lower cutoff voltage, the exemplary lithium cobalt oxide material doped with zirconium demonstrates slightly higher coulombic efficiency (CE) as shown in Table 8. To be more specific, under the upper cutoff voltage of 4.45 V, the coulombic efficiency (CE) is 96.3%, while the coulombic efficiency (CE) is 92.9% under the upper cutoff voltage of 4.5 V.

TABLE-US-00008 TABLE 8 Measured electric performance of lithium-ion-battery cells made from exemplary LCO doped with Zirconium materials Example #16 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.45 187.297 180.337 96.3 4.5 195.108 181.306 92.9

[0160] In other embodiments, the compositions and formulations of the present inventions being tested are as shown in the below Table 8. The compositions of the present inventions, prepared according to Example #22-#26, have a chemical formula of $\text{Li}_{x}\text{Co}_{y}\text{O}_{z}\text{Math.doped Mg}_{a}\text{Al}_{b}\text{Zr}_{c}$, wherein a ratio of $x:y:a:b:c$ is equivalent to

$\text{M}_{\text{LiSalt}}:\text{M}_{\text{CoSalt}}:\text{M}_{\text{MgSalt}}:\text{M}_{\text{AlSalt}}:\text{M}_{\text{ZrSalt}}$, wherein x is 1.0, y is 1.0, the a is from 0-0.05 ($0 \leq a \leq 0.05$), the b is from 0-0.05 ($0 \leq b \leq 0.05$), the c is from 0-0.05 ($0 \leq c \leq 0.05$).

[0161] In Example #22-#26, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO_{3}) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO)_{3}) and combinations thereof, exemplary first metal-containing salt Me1 include, but not limited to, magnesium nitrate (Mg(NO)_{3}) and combinations thereof, exemplary second metal-containing salt Me2 compound include, but not limited to, aluminum nitrate (Al(NO)_{3}) and combinations thereof, exemplary third metal-containing salt Me3 compound include, but not limited to, zirconium nitrate (Zr(NO)_{3}) and combinations thereof. The annealing temperature and annealing time in the Example #22-#26 were heated to 1020° C. for 17 hours. The list of chemistries used for in the present invention is displayed in Table 9.

TABLE-US-00009 TABLE 9 Exemplary compositions of measured LCO material doped with one or more metal dopants Example M_{LiSalt} M_{CoSalt} M_{MgSalt} M_{AlSalt} M_{ZrSalt} Anneal Temp Anneal Time # M_{LiSalt} M_{CoSalt} (Mg(NO)_{3}) (Al(NO)_{3}) (Zr(NO)_{3}) (° C.) (hour) 22 1.0 1.0 0.0017 0 0 1020 17 23 1.0 1.0 0.0045 0 0 1020 17 24 1.0 1.0 0 0.0045 0 1020 17 25 1.0 1.0 0.0010 0.0035 0 1020 17 26 1.0 1.0 0 0 0.0017 1020 17

[0162] Table 10 illustrates testing results of exemplary measured LCO material doped with one or more metal dopants compositions (Example #22-#26). One observation is that the testing results of the ratio of the measured LCO material compositions of Li:Co:Mg:Al:Zr are within an expected range from the prepared molar ratio of

$\text{M}_{\text{LiSalt}}:\text{M}_{\text{CoSalt}}:\text{M}_{\text{MgSalt}}:\text{M}_{\text{AlSalt}}:\text{M}_{\text{ZrSalt}}$ being prepared.

TABLE-US-00010 TABLE 10 Exemplary compositions of measured LCO material doped with one or more metal dopants Example # Li Ni Co Al Mg Zr 22 1.0040 0.0014 0.998 0.0003 0.0021 0.0002 23 1.0099 0 0.999 0.0007 0.0053 0 24 1.0131 0 0.9954 0.0044 0.0001 0 25 1.0165 0 0.9954 0.0044 0.0017 0 26 1.0140 0 0.9995 0.0003 0.0001 0.0009

[0163] Table 11 illustrates testing results of tap density (TD) and contaminants of crystalized lithium cobalt oxide materials doped with one or more metal dopants after annealing process of exemplary LCO doped with one or more metal dopants (Example #22-26). To obtain an ideal lithium cobalt oxide material with high discharge capacity, excellent cycling performance and high-volume energy density, the morphology and tap density of the material have to be controlled precisely during the preparation process. It is found that the tap density of the obtained precursor is around 2.38 (g/cc), which can be attributed to the homogeneous distributions of particles with good packing properties.

TABLE-US-00011 TABLE 11 Measurement of tap density (TD) & contaminants of exemplary LCO materials doped with one or more metal dopants Example # TD (g/cc) Li₂CO₃ LiOH 22 2.17 0.024 0 23 2.38 0.036 0.005 24 2.27 0.035 0.005 25 2.18 0.027 0 26 2.10 0.031 0

[0164] Table 12 illustrates testing results of electric capacity and coulombic efficiency (CE) of

battery cells made by lithium cobalt oxide materials doped with magnesium tested under different upper cutoff voltages from 4.45 voltage to 4.6 voltage, prepared according to Example #22. One observation can be found that the battery samples made by exemplary lithium cobalt oxide materials doped with magnesium overall show a high coulombic efficiency (CE) under different cutoff voltages. For example, under the upper cutoff voltage of 4.45 V, the discharge capacity and the coulombic efficiency (CE) is around 181.404 mAh/g and 98%, respectively. In another example, under the upper cutoff voltage of 4.5 V, the discharge capacity and coulombic efficiency (CE) is around 194.346 mAh/g and 99.4%, respectively. In still another example, under the upper cutoff voltage of 4.6 V, the discharge capacity and the coulombic efficiency (CE) is around 226.019 mAh/g and 97%, respectively.

[0165] Referring back to Table 8, further observation can be found that samples of batter cells made from lithium cobalt oxide materials doped with magnesium (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0017) demonstrate higher coulombic efficiency (CE) than the CE made from lithium cobalt oxide materials doped with zirconium (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Zr.sub.0.0017). To be more specific, the measured CE of battery cells made from exemplary LCO doped with zirconium materials is ranged from 92.9% to 96.3%, while the measured CE of battery cells made from exemplary LCO doped with magnesium materials is ranged from 96.9% to 99.4%

TABLE-US-00012 TABLE 12 Measured electric performance of lithium-ion-battery cells made from exemplary LCO doped with Magnesium materials Example # 22 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.45 185.104 181.404 98 4.45 185.557 181.493 97.8 4.5 195.423 194.346 99.4 4.5 193.576 191.201 98.8 4.6 233.032 226.019 97 4.6 233.108 225.913 96.9

[0166] Table 13 illustrates testing results of electric capacity and coulombic efficiency (CE) of battery cells made by lithium cobalt oxide materials doped with magnesium tested under different upper cutoff voltages from 4.3 voltage to 4.6 voltage, prepared according to Example #23. One observation can be found that the battery samples made by exemplary lithium cobalt oxide materials doped with magnesium overall show a high coulombic efficiency (CE) under different cutoff voltages. For example, under the upper cutoff voltage of 4.3 V, the discharge capacity and the coulombic efficiency (CE) is around 160.434 mAh/g and 97.7%, respectively. In another example, under the upper cutoff voltage of 4.45 V, the discharge capacity and the coulombic efficiency (CE) is around 183.173 mAh/g and 97.6%, respectively. In still another example, under the upper cutoff voltage of 4.5 V, the discharge capacity and the coulombic efficiency (CE) is around and 193.217 mAh/g and 97.3%, respectively. In yet another example, under the upper cutoff voltage of 4.6 V, the discharge capacity and the coulombic efficiency (CE) is around 228.309 mAh/g and 96.6%, respectively

[0167] Referring back to Table 12, further observation can be found that samples of batter cells made from lithium cobalt oxide materials doped with different percentage of magnesium demonstrate similar coulombic efficiency (CE). To be more specific, the measured CE of battery cells made from exemplary LCO doped with 0.0017 magnesium materials (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0017) is ranged from 96.9% to 99.4% under different upper cutoff voltage ranged from 4.45 V to 4.6 V, while the measured CE of battery cells made from exemplary LCO doped with 0.0045 (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0045) magnesium materials is ranged from 95.7% to 97.7% under different upper cutoff voltage ranged from 4.3 V to 4.6 V

TABLE-US-00013 TABLE 13 Measured electric performance of lithium-ion-battery cells made from exemplary LCO doped Magnesium materials Example # 23 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.3 164.16 160.434 97.7 4.3 166.873 159.676 95.7 4.45 187.556 182.718 97.4 4.45 187.757 183.173 97.6 4.5 198.574 193.217 97.3 4.6 236.302 228.309 96.6

[0168] Table 14 illustrates testing results of electric capacity and coulombic efficiency (CE) of battery cells made by lithium cobalt oxide materials doped with aluminum tested under different upper cutoff voltages from 4.4 voltage to 4.6 voltage, prepared according to Example #24. One observation can be found that the cutoff voltage the battery samples made by exemplary lithium cobalt oxide materials doped with aluminum overall have a high coulombic efficiency (CE). For example, under the upper cutoff voltage of 4.4 V, the discharge capacity and the coulombic efficiency (CE) is around 173.592 mAh/g and 97%, respectively. In another example, the discharge capacity and under the upper cutoff voltage of 4.45V, the coulombic efficiency (CE) is around 181.861 mAh/g and 97.3%, respectively. In still another example, under the upper cutoff voltage of 4.5 V, the discharge capacity and the coulombic efficiency (CE) is around 191.712 mAh/g and 97%, respectively. In yet another example, under the upper cutoff voltage of 4.6 V, the discharge capacity and the coulombic efficiency (CE) is around 225.187 mAh/g and 96.4%.

[0169] Referring back to Table 13, further observation can be found that samples of batter cells made from lithium cobalt oxide materials doped with magnesium (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0045) demonstrate similar coulombic efficiency (CE) to lithium cobalt oxide materials doped with aluminum (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Al.sub.0.0045). To be more specific, the measured CE of battery cells made from exemplary LCO doped with magnesium materials is ranged from 95.7% to 97.7%, while the measured CE of battery cells made from exemplary LCO doped with aluminum materials is ranged from 96.4% to 97.3%

TABLE-US-00014 TABLE 14 Measured electric performance of lithium-ion-battery cells made from exemplary LCO doped Aluminum materials Example # 24 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.4 178.991 173.592 97 4.45 186.964 181.861 97.3 4.45 186.955 181.729 97.2 4.5 197.624 191.712 97 4.6 233.659 225.187 96.4

[0170] Table 15 illustrates testing results of electric capacity and coulombic efficiency (CE) of battery cells made by lithium cobalt oxide materials doped with aluminum tested under different upper cutoff voltages from 4.45 voltage to 4.6 voltage, prepared according to Example #25. One observation can be found that the cutoff voltage the battery samples made by exemplary lithium cobalt oxide materials doped with magnesium and aluminum overall have a high coulombic efficiency (CE). For example, under the upper cutoff voltage of 4.45 V, the discharge capacity and the coulombic efficiency (CE) is around 180.661 mAh/g and 93.2%, respectively. In another example, under the upper cutoff voltage of 4.5V, the discharge capacity and the coulombic efficiency (CE) is around 190.324 mAh/g and 96.5%, respectively. In still another example, under the upper cutoff voltage of 4.6 V, the discharge capacity and the coulombic efficiency (CE) is around 224.68 mAh/g and 95.8%, respectively.

[0171] Referring back to Table 11, Table 12, Table 13 and Table 14, further observation can be found that samples of battery cells made from lithium cobalt oxide materials with one metal dopant in average demonstrate higher coulombic efficiency (CE) to lithium cobalt oxide materials with two metal dopants (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0010Al.sub.0.0035). To be more specific, the measured CE of battery cells made from exemplary LCO with one metal dopant is ranged from 96.4% to 99.4%, while the measured CE of battery cells made from exemplary LCO doped with magnesium and aluminum materials is ranged from 93.2% to 96.5%

TABLE-US-00015 TABLE 15 Measured electric performance of lithium-ion-battery cells made from exemplary LCO • doped Mg, Al materials Example # 25 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.45 193.756 180.661 93.2 4.5 197.212 190.324 96.5 4.6 232.562 222.25 95.6 4.6 234.493 224.638 95.8

[0172] Table 16 illustrates testing results of electric capacity and coulombic efficiency (CE) of battery cells made by lithium cobalt oxide materials doped with zirconium tested under different upper cutoff voltages from 4.3 voltage to 4.6 voltage, prepared according to Example #26. One

observation can be found that the cutoff voltage the battery samples made by exemplary lithium cobalt oxide materials doped with zirconium overall show a high coulombic efficiency (CE). For example, under the upper cutoff voltage of 4.3 V, the discharge capacity and the coulombic efficiency (CE) is around 159.025 mAh/g and 98.3%, respectively. In another example, under the upper cutoff voltage of 4.45 V, the discharge capacity and coulombic efficiency (CE) is around 181.71 mAh/g and 98.0%, respectively. In still another example, under the upper cutoff voltage of 4.5 V, the discharge capacity and the coulombic efficiency (CE) is around 194.107 mAh/g and 97.9%, respectively. In yet another example, under the upper cutoff voltage of 4.6 V, the discharge capacity and the coulombic efficiency (CE) is around 230.09 mAh/g and 96.9%, respectively.

[0173] Referring back to Table 8, further observation can be found that samples of batter cells made from lithium cobalt oxide materials doped with zirconium annealed at 950° C. for 17 hours (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Zr.sub.0.0017) demonstrate lower coulombic efficiency (CE) than the CE made from lithium cobalt oxide materials doped with zirconium annealed at 1020° C. for 17 hours (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Zr.sub.0.0017). To be more specific, the measured CE of battery cells made from exemplary LCO doped with zirconium materials annealed at 1020° C. is ranged from 96.9% to 98.3%, while the measured CE of battery cells made from exemplary LCO doped with zirconium materials annealed at 950° C. is ranged from 92.9% to 96.3%

TABLE-US-00016 TABLE 16 Measured electric performance of lithium-ion-battery cells made from exemplary LCO • doped Zr materials Example # 26 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.3 161.831 159.025 98.3 4.45 185.369 181.71 98.0 4.5 198.267 194.107 97.9 4.6 237.538 230.09 96.9

[0174] In still other embodiments, the compositions and formulations of the present inventions being tested are as shown in the below Table 17. The compositions of the present inventions for Example 32 having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, wherein a ratio of x:y:a is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt, wherein x is 1.0, y is 1.0, a is 0.0017.

[0175] In Example 32, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO.sub.3) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO.sub.3).sub.2) and combinations thereof, exemplary first metal-containing salt Me1 include, but not limited to, magnesium nitrate (Mg(NO.sub.3).sub.2) and combinations thereof. The annealing temperature and annealing time in the Example 32 were heated to 1090° C. for 17 hours. The list of chemistries used for in the present invention is displayed in Table 17.

TABLE-US-00017 TABLE 17 Exemplary compositions of measured LCO material doped with magnesium Example M.sub.Me1Salt M.sub.Me2Salt M.sub.Me3Salt Anneal Temp Anneal Time # M.sub.LiSalt M.sub.CoSalt (Mg(NO.sub.3).sub.2) (Al(NO.sub.3).sub.3) (Zr(NO.sub.3).sub.4) (° C.) (hour) 32 1.0 1.0 0.0017 0 0 1090 17

[0176] Table 18 illustrates testing results of exemplary measured LCO material compositions (Example #32). One observation is that the testing results of the ratio of the measured LCO material compositions of Li:Co:Mg are within an expected range from the prepared molar ratio of M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt being prepared.

TABLE-US-00018 TABLE 18 Exemplary compositions of measured LCO material doped with magnesium Example # Li Ni Co Al Mg Zr 32 1.0072 0 0.9979 0.0018 0.0020 0

[0177] Table 19 illustrates testing results of electric capacity and coulombic efficiency (CE) of battery cells made by lithium cobalt oxide materials doped with magnesium annealed at 1090° C. tested under different upper cutoff voltages from 4.3 voltage to 4.6 voltage, prepared according to Example #32. One observation can be found that the battery samples made by exemplary lithium cobalt oxide materials doped with magnesium overall show a high coulombic efficiency (CE) under different cutoff voltages. For example, under the upper cutoff voltage of 4.3 V, the discharge

capacity and the coulombic efficiency (CE) is around 146.55 mAh/g and 98.2%, respectively. In another example, under the upper cutoff voltage of 4.45 V, the discharge capacity and the coulombic efficiency (CE) is around 175.842 mAh/g and 98.3%, respectively. In still another example, under the upper cutoff voltage of 4.5 V, the discharge capacity and the coulombic efficiency (CE) is around 175.7 mAh/g and 98.2%, respectively. In yet another example, under the upper cutoff voltage of 4.6 V, the discharge capacity and the coulombic efficiency (CE) is around 217.2 mAh/g and 97.5%, respectively

[0178] Referring back to Table 12, further observation can be found that samples of battery cells made from lithium cobalt oxide materials doped with magnesium

(Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0017) annealed at different annealing temperatures demonstrate similar coulombic efficiency (CE). To be more specific, the measured CE of battery cells made from exemplary LCO doped with magnesium materials annealed at 1020° C. is ranged from 96.9% to 99.4%, while the measured CE of battery cells made from exemplary LCO doped with magnesium materials annealed at 1090° C. is ranged from 97.1% to 98.3%

TABLE-US-00019 TABLE 19 Measured electric performance of lithium-ion-battery cells made from exemplary LCO • doped Mg materials Example # 32 Upper cut off 1st charge 1st discharge voltage, V capacity, mAh/g capacity, mAh/g 1st CE, % 4.3 149.23 146.55 98.2 4.45 178.845 175.842 98.3 4.5 178.9 175.7 98.2 4.5 192.3 186.7 97.1 4.6 222.8 217.2 97.5

[0179] FIG. 4 illustrates testing results of the discharge profile of electric capacity of lithium ion batteries prepared from lithium cobalt oxide materials doped with one or metal dopants of the invention. In one embodiment, line 410 represents lithium cobalt oxide materials doped with zirconium (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Zr.sub.0.0017). In another embodiment, line 430 represents lithium cobalt oxide materials doped with Aluminum

(Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Al.sub.0.0045). All dopants affect the discharge capacity of the lithium cobalt oxide materials. One observation can be found that the samples LCO dependent on different dopant composition ratios has a higher capacity at a higher voltage as shown in FIG. 4.

[0180] Other observation can be found that at the same upper cut-off voltage, the discharge capacities drop slightly due to the different dopants at different ratios, however no significant drop can be observed, which confirms that substituted dopant levels do not affect the electrochemical performance of the sample significantly.

[0181] Further observation can be found that at the upper cut-off voltage 4.6 V, the lithium cobalt oxide materials doped with zirconium (Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Zr.sub.0.0017) have the optimal and the highest discharge capacity of 230.09 mAh/g and 96.7% coulombic efficiency (CE) among other composition ratios of cathode material mixtures as shown in FIG. 4.

[0182] FIG. 5A, FIG. 5B and FIG. 5C illustrate the discharge profile of electric capacity of lithium ion batteries at different cut-off voltages (from 4.45 voltage to 4.6 voltage), where the lithium ion batteries are prepared according to Example #22-#26 of lithium cobalt oxide materials doped with one or more metal dopants of the invention.

[0183] FIG. 5A is a column graph illustrating the discharge profile of electric capacity of lithium ion batteries prepared according to Example #22-#26 of the invention at 4.45 cut-off voltage. One observation can be found that at 4.45 cut-off voltage, the discharge capacities drop slightly due to the different dopants at different ratios. Further observation can be found that at the upper cut-off voltage 4.45 V, the lithium cobalt oxide materials doped with magnesium

(Li.sub.1.0Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.0045), prepared according to Example #23, have the highest discharge capacity of 183.17 mAh/g among other composition ratios of Example #22-26 as shown in FIG. 5A.

[0184] FIG. 5B is a column graph illustrating the discharge profile of electric capacity of lithium ion batteries prepared from Example #22-26 of the invention at 4.5 voltage. One observation can be found that at 4.5 cut-off voltage, the discharge capacities drop slightly due to the different dopants

at different ratios. Further observation can be found that at the upper cut-off voltage 4.5 V, the lithium cobalt oxide materials doped with magnesium ($\text{Li}_{1.0}\text{Co}_{1.0}\text{O}_{2.0}\text{Math.doped Mg}_{0.0017}$), prepared according to Example #22, and lithium cobalt oxide materials doped with zirconium ($\text{Li}_{1.0}\text{Co}_{1.0}\text{O}_{2.0}\text{Math.doped Zr}_{0.0017}$), prepared according to Example #26, have higher discharge capacity of 194.35 mAh/g and 194.11 mAh/g, respectively, among other composition ratios of Example #22-26 as shown in FIG. 5B.

[0185] FIG. 5C is a column graph illustrating the discharge profile of electric capacity of lithium ion batteries prepared from Example #22-26 of the invention at 4.6 voltage. As shown in FIG. 5C, a higher discharge capacity reading is observed as the cut-off voltage increases. Further observation can be found that the lithium cobalt oxide materials doped with zirconium ($\text{Li}_{1.0}\text{Co}_{1.0}\text{O}_{2.0}\text{Math.doped Zr}_{0.0017}$), prepared according to Example #26, have the higher discharge capacity of 230.09 mAh/g among other composition ratios of Example #22-26.

[0186] FIG. 6 is a graph illustrating cycling performance of samples of battery cells made from lithium cobalt oxide materials doped with different metal dopants and concentrations at a cutoff voltage of 4.5 V. FIG. 6 compares the cycle and discharge capability test results for the doped examples.

[0187] In one embodiment, line 610 illustrates the charge cycles of battery cells made from lithium cobalt oxide doped with zirconium ($\text{LiCoO}_{2.0}\text{Math.doped Zr}_{0.17\%}$). In another embodiment, line 620 illustrates the charge cycles of battery cells made from lithium cobalt oxide doped with magnesium ($\text{LiCoO}_{2.0}\text{Math.doped Mg}_{0.17\%}$). In yet another embodiment, line 630 illustrates the charge cycles of battery cells made from lithium cobalt oxide doped with magnesium ($\text{LiCoO}_{2.0}\text{Math.doped Mg}_{0.45\%}$). In yet another embodiment, line 640 illustrates the charge cycles of battery cells made from lithium cobalt oxide doped with magnesium and aluminum ($\text{LiCoO}_{2.0}\text{Math.doped Mg}_{0.1\%}, \text{Al}_{0.35\%}$). In yet another embodiment, line 650 illustrates the charge cycles of battery cells made from lithium cobalt oxide doped with aluminum ($\text{LiCoO}_{2.0}\text{Math.doped Al}_{0.45\%}$).

[0188] Observation can be seen by FIG. 6 is that the cycling of battery cells doped with zirconium containing compound ($\text{LiCoO}_{2.0}\text{Math.doped Zr}_{0.17\%}$) performs better than the battery cells doped with other metals containing compound. Further illustration can be seen from FIG. 6 is that between 0-25 battery cycles, the battery cell made from lithium cobalt oxide materials doped with zirconium ($\text{LiCoO}_{2.0}\text{Math.doped Zr}_{0.17\%}$) has higher capacity than the other samples made from lithium cobalt oxide materials doped with magnesium, lithium cobalt oxide materials doped with aluminum, and lithium cobalt oxide materials doped with magnesium and aluminum of the battery cells. Further observation is that the discharge capacity of these samples starts to slowly fade by the time they reach 25th cycle.

[0189] FIG. 7A and FIG. 7B are scanning electron microscopy (SEM) images of one example of crystalized lithium cobalt oxide materials doped with magnesium ($\text{LiCoO}_{2.0}\text{Math.doped Mg}_{0.17\%}$) of the invention after the annealing process at 1020° C. for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt is $\text{M}_{\text{sub}}.\text{LiSalt}:\text{M}_{\text{sub}}.\text{CoSalt}:\text{M}_{\text{sub}}.\text{MgSalt}$. The present invention having a chemical formula of $\text{Li}_{\text{sub}.x}\text{Co}_{\text{sub}.y}\text{O}_{\text{sub}.z}\text{Math.doped Mg}_{\text{sub}.a}$, wherein a ratio of x:y:a is equivalent to $\text{M}_{\text{sub}}.\text{LiSalt}:\text{M}_{\text{sub}}.\text{CoSalt}:\text{M}_{\text{sub}}.\text{MgSalt}$, wherein the x is 1.0, the y is 1.0, and a is 0.0017. In FIG. 7A and FIG. 7B, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate ($\text{LiNO}_{3.0}$) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate ($\text{Co}(\text{NO}_{3.0})_{2.0}$) and combinations thereof, exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate ($\text{Mg}(\text{NO}_{3.0})_{2.0}$) and combinations thereof.

[0190] FIG. 7A illustrates the morphology and particle size of one example of lithium cobalt oxide

material doped with magnesium particles at an annealing temperature of 1020° C. for 17 hours having crystalized structure. In addition, FIG. 7B shows a closer look of FIG. 7A. In one example as shown in FIG. 7B one lithium cobalt oxide material doped with magnesium particle **710** has a crystal structure.

[0191] FIG. 7C and FIG. 7D are scanning electron microscopy (SEM) images of another example of solid particles of an oxide material doped with magnesium (LiCoO₂.Math.doped Mg.sub.0.17%) after a drying process inside a drying chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, wherein a ratio of x:y:a is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt, wherein the x is 1.0, the y is 1.0, and a is 0.0017. In FIG. 7C and FIG. 7D, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO₃) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO₃)₂) and combinations thereof, exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO₃)₂) and combinations thereof.

[0192] FIG. 7C illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium particles after the drying process having crystalized structure. In addition, FIG. 7D shows a closer look of FIG. 7C. In one example as shown in FIG. 7D, one solid particle of a lithium cobalt oxide material doped with magnesium **720** is spherical in shape.

[0193] FIG. 8A and FIG. 8B are scanning electron microscopy (SEM) images of crystalized lithium cobalt oxide materials doped with magnesium (LiCoO₂.Math.doped Mg.sub.0.45%) of the invention after the annealing process at 1020° C. for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, wherein a ratio of x:y:a is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt, Wherein the x is 1.0, the y is 1.0, and a is 0.0045. In FIG. 8A and FIG. 8B, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO₃) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO₃)₂) and combinations thereof, exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO₃)₂) and combinations thereof.

[0194] FIG. 8A illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium particles at an annealing temperature of 1020° C. for 17 hours having crystalized structure. In addition, FIG. 8B shows a closer look of FIG. 8A. In one example as shown in FIG. 8B one lithium cobalt oxide material doped with magnesium particle **810** has a crystal structure. Referring back to FIG. 7B, further observation can be found that the morphology does not change significantly with different percentage of magnesium dopants.

[0195] FIG. 8C and FIG. 8D are scanning electron microscopy (SEM) images of another example of solid particles of an oxide material doped with magnesium (LiCoO₂.Math.doped Mg.sub.0.45%) after a drying process inside a drying chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, wherein a ratio of x:y:a is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt, wherein the x is 1.0, the y is 1.0, and a is 0.0045. In FIG. 8C and FIG. 8D, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO₃) and combinations thereof, exemplary cobalt-containing salt compound

include, but not limited to cobalt nitrate (Co(NO.sub.3).sub.2) and combinations thereof, exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO.sub.3).sub.2) and combinations thereof.

[0196] FIG. 8C illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium particles after the drying process having crystalized structure. In addition, FIG. 8D shows a closer look of FIG. 8C. In one example as shown in FIG. 8D, one solid particle of a lithium cobalt oxide material doped with magnesium **820** is spherical in shape.

[0197] FIG. 9A and FIG. 9B are scanning electron microscopy (SEM) images of crystalized lithium cobalt oxide materials doped with aluminum (LiCoO.sub.2.Math.doped Al.sub.0.45%) of the invention after the annealing process at 1020° C. for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, an aluminum-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.AlSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Al.sub.b, wherein a ratio of x:y:b is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.AlSalt, wherein the x is 1.0, the y is 1.0, and b is 0.0045. In FIG. 9A and FIG. 9B, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO.sub.3) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO.sub.3).sub.2) and combinations thereof, exemplary aluminum-containing salt (Al) include, but not limited to, aluminum nitrate (Al(NO.sub.3).sub.3) and combinations thereof.

[0198] FIG. 9A illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with aluminum particles at an annealing temperature of 1020° C. for 17 hours having crystalized structure. In addition, FIG. 9B shows a closer look of FIG. 9A. In one example as shown in FIG. 9B one lithium cobalt oxide material doped with aluminum particle **910** has a crystal structure.

[0199] FIG. 9C and FIG. 9D are scanning electron microscopy (SEM) images of another example of solid particles of an oxide material doped with aluminum (LiCoO.sub.2.Math.doped Al.sub.0.45%) after a drying process inside a drying chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, an aluminum-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.AlSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Al.sub.b, wherein a ratio of x:y:b is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.AlSalt, wherein the x is 1.0, the y is 1.0, and b is 0.0045. In FIG. 9C and FIG. 9D, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO.sub.3) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO.sub.3).sub.2) and combinations thereof, exemplary aluminum-containing salt (Al) include, but not limited to, aluminum nitrate (Al(NO.sub.3).sub.3) and combinations thereof.

[0200] FIG. 9C illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with aluminum particles after the drying process having crystalized structure. In addition, FIG. 9D shows a closer look of FIG. 9C. In one example as shown in FIG. 9D, one solid particle of a lithium cobalt oxide material doped with aluminum **920** is spherical in shape.

[0201] FIG. 10A and FIG. 10B are scanning electron microscopy (SEM) images of crystalized lithium cobalt oxide materials doped with magnesium and aluminum (LiCoO.sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%) of the invention after the annealing process at 1020° C. for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt, an aluminum-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, Al.sub.b, wherein a ratio of x:y:a:b is

equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt, wherein the x is 1.0, the y is 1.0, a is 0.001 and b is 0.0035. In FIG. **10A** and FIG. **10B**, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO₃) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO₃)₂) and combinations thereof. Exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO₃)₂) and combinations thereof. Exemplary aluminum-containing salt (Al) include, but not limited to, aluminum nitrate (Al(NO₃)₃) and combinations thereof.

[0202] FIG. **10A** illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium and aluminum particles at an annealing temperature of 1020° C. for 17 hours having crystalized structure. In addition, FIG. **10B** shows a closer look of FIG. **10A**. In one example as shown in FIG. **10B** one lithium cobalt oxide material doped with magnesium and aluminum particle **1010** has a crystal structure.

[0203] FIG. **10C** and FIG. **10D** are scanning electron microscopy (SEM) images of another example of solid particles of an oxide material doped with magnesium and aluminum (LiCoO₂.Math.doped Mg.sub.0.10%, Al.sub.0.35%) after a drying process inside a drying chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt, an aluminum-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt. The present invention having a chemical formula of Li_xCo_yO_z.Math.doped Mg_a, Al_b, wherein a ratio of x:y:a:b is equivalent to

M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt, wherein the x is 1.0, the y is 1.0, a is 0.001 and b is 0.0035. In FIG. **10C** and FIG. **10D**, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO₃) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO₃)₂) and combinations thereof. Exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO₃)₂) and combinations thereof. Exemplary aluminum-containing salt (Al) include, but not limited to, aluminum nitrate (Al(NO₃)₃) and combinations thereof.

[0204] FIG. **10C** illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium and aluminum particles after the drying process having crystalized structure. In addition, FIG. **10D** shows a closer look of FIG. **10C**. In one example as shown in FIG. **10D**, one solid particle of a lithium cobalt oxide material doped with magnesium and aluminum **1020** is spherical in shape.

[0205] FIG. **11A** and FIG. **11B** are scanning electron microscopy (SEM) images of crystalized lithium cobalt oxide materials doped with zirconium (LiCoO₂.Math.doped Zr.sub.0.17%) of the invention after the annealing process at 1020° C. for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.ZrSalt. The present invention having a chemical formula of Li_xCo_yO_z.Math.doped Zr_c, wherein a ratio of x:y:c is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.ZrSalt, wherein the x is 1.0, the y is 1.0, and c is 0.0017. In FIG. **11A** and FIG. **11B**, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO₃) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO₃)₂) and combinations thereof, and exemplary zirconium-containing salt (Zr) include, but not limited to, zirconium nitrate (Zr(NO₃)₄) and combinations thereof.

[0206] FIG. **11A** illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with zirconium particles at an annealing temperature of 1020° C. for 17 hours having crystalized structure. In addition, FIG. **11B** shows a closer look of FIG. **11A**. In one

example as shown in FIG. 11B one lithium cobalt oxide material doped with zirconium particle **1110** has a crystal structure. Referring back to FIG. 7B-FIG. 10B, further observation can be found that the morphology does not change significantly with different dopants.

[0207] FIG. 12A and FIG. 12B are scanning electron microscopy (SEM) images of one example of crystalized lithium cobalt oxide materials doped with magnesium (Li.sub.0.97Co.sub.1.0O.sub.2.Math.doped Mg.sub.0.17%) of the invention after the annealing process at 1020° C. for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, wherein a ratio of x:y:a is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt, wherein the x is 0.97, the y is 1.0, and a is 0.0017. In FIG. 12A and FIG. 12B, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO.sub.3) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO.sub.3).sub.2) and combinations thereof, exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO.sub.3).sub.2) and combinations thereof.

[0208] FIG. 12A illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium particles at an annealing temperature of 1020° C. for 17 hours having crystalized structure. In addition, FIG. 12A shows a closer look of FIG. 12A. In one example as shown in FIG. 12B one lithium cobalt oxide material doped with magnesium particle **1210** has a crystal structure. Referring back to FIG. 7B-FIG. 11B, further observation can be found that the morphology does not change significantly with different dopants.

[0209] FIG. 13A and FIG. 13B are scanning electron microscopy (SEM) images of one example of solid particles of an oxide material doped with magnesium and aluminum (LiCoO.sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%) after a drying process inside a drying chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt, an aluminum-containing salt is M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt. The present invention having a chemical formula of Li.sub.xCo.sub.yO.sub.z.Math.doped Mg.sub.a, Al.sub.b, wherein a ratio of x:y:a:b is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt, wherein the x is 1.0, the y is 1.0, a is 0.001 and b is 0.0035.

[0210] In FIG. 13A and FIG. 13B, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO.sub.3) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate (Co(NO.sub.3).sub.2) and combinations thereof. Exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate (Mg(NO.sub.3).sub.2) and combinations thereof. Exemplary aluminum-containing salt (Al) include, but not limited to, aluminum nitrate (Al(NO.sub.3).sub.3) and combinations thereof.

[0211] FIG. 13A illustrates the morphology and particle size of one example of lithium cobalt oxide material doped with magnesium and aluminum particles after the drying process having crystalized structure. In addition, FIG. 13B shows a closer look of FIG. 13A. In one example as shown in FIG. 13B, one solid particle of a lithium cobalt oxide material doped with magnesium and aluminum **1310** is spherical in shape.

[0212] Referring back to FIG. 13A and FIG. 13B, the tap density (“TD”) and particle size distribution (known as “PSD” or “SPAN”) of the solid particles of a lithium cobalt oxide material is shown as Table 20.

[0213] To be precise, the “SPAN” value represents a degree of particle size distribution (“PSD”), defined as (D90–D10)/D50. “D10”, “D50”, and “D90” are defined as the particle size at 10%, 50%, and 90% of the cumulative volume % distribution. D50 represents an average particle size

which is larger than 50% (by number) of the total particles that are present (also known as the median diameter). D90 and D10 respectively stand for particle diameters that are larger than 90% and 10% of all the particles (by number). Further, the particle size is expressed in μm . The tapped density of the solid particles of a lithium cobalt oxide material is 0.86 g/cc, D10 value is 3.69 μm , D50 value is 7.71 μm , D90 value is 13.3 μm and PSD value is 1.25.

TABLE-US-00020 TABLE 20 Measurement of tap density (TD) and particle size of solid particles of a lithium cobalt oxide material after a drying process Tap Density D10 D50 D90 PSD 0.86 3.69 7.71 13.3 1.25

[0214] FIG. 14A-14C, FIG. 15A-15C, FIG. 16A-16C, and FIG. 17A-17C are scanning electron microscopy (SEM) images of crystalized lithium cobalt oxide materials doped with magnesium and aluminum (LiCoO_2 .sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%) of the invention after the annealing process at different annealing temperatures (from 900° C. to 1050° C.) for 17 hours inside the annealing chamber. The SEM Image shows the compositions and formulations of the present inventions having a molar ratio of a lithium-containing salt, a cobalt-containing salt, a magnesium-containing salt, an aluminum-containing salt is

M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt. The present invention having a chemical formula of $\text{Li}_x\text{Co}_y\text{O}_z$.sub.2.Math.doped Mg.sub.a, Al.sub.b, wherein a ratio of x:y:a:b is equivalent to M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt, wherein the x is 1.0, the y is 1.0, a is 0.001 and b is 0.0035.

[0215] In FIG. 14A-14C, FIG. 15A-15C, FIG. 16A-16C, and FIG. 17A-17C, exemplary lithium-containing salt compound include, but not limited to, lithium nitrate (LiNO_3 .sub.3) and combinations thereof, exemplary cobalt-containing salt compound include, but not limited to cobalt nitrate ($\text{Co(NO}_3)_2$.sub.2) and combinations thereof. Exemplary magnesium-containing salt (Mg) include, but not limited to, magnesium nitrate ($\text{Mg(NO}_3)_2$.sub.2) and combinations thereof. Exemplary aluminum-containing salt (Al) include, but not limited to, aluminum nitrate ($\text{Al(NO}_3)_3$.sub.3) and combinations thereof.

[0216] In one example as shown in FIG. 14A, the SEM shows the morphology and particle size of lithium cobalt oxide material particles doped with magnesium and aluminum (LiCoO_2 .sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%) at an annealing temperature of 900° C. for 17 hours having crystalized structure. FIG. 14B and FIG. 14C show a closer look of FIG. 14A.

[0217] In another example as shown in FIG. 15A, the SEM shows the morphology and particle size of lithium cobalt oxide material particles doped with magnesium and aluminum at an annealing temperature of 950° C. for 17 hours having crystalized structure. FIG. 15B and FIG. 15C show a closer look of FIG. 15A.

[0218] Referring back to FIG. 16A, the SEM shows the morphology and particle size of yet another example of lithium cobalt oxide material particles doped with magnesium and aluminum (LiCoO_2 .sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%) at an annealing temperature of 1000° C. for 17 hours having crystalized structure. FIG. 16B and FIG. 16C show a closer look of FIG. 16A.

[0219] In still another example as shown in FIG. 17A, the SEM shows the morphology and particle size of lithium cobalt oxide material particles doped with magnesium and aluminum (LiCoO_2 .sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%) at an annealing temperature of 1050° C. for 17 hours having crystalized structure. FIG. 17B and FIG. 17C show a closer look of FIG. 17A.

[0220] Referring back to FIG. 13-17, Table 21 illustrates testing results of samples of exemplary measured LCO material doped with magnesium and aluminum (LiCoO_2 .sub.2.Math.doped Mg.sub.0.10%, Al.sub.0.35%). One observation is that the testing results of the ratio of the measured LCO material compositions of Li:Co:Mg:Al are within an expected range from the prepared molar ratio of M.sub.LiSalt:M.sub.CoSalt:M.sub.MgSalt:M.sub.AlSalt being prepared. TABLE-US-00021 TABLE 21 Exemplary measured LCO material compositions Corresponding FIG. Li Co Al Mg Condition FIG. 13A, 1.0862 0.9875 0.0125 0.0040 After Drying FIG. 13B FIG. 14A, 1.0629 0.9873 0.0127 0.0041 Anneal at FIG. 14B 900° C. FIG. 15A, 1.0424 0.9862 0.0138

0.0041 Anneal at FIG. 15B 950° C. FIG. 16A, 1.0313 0.9867 0.0133 0.0040 Anneal at FIG. 16B 1000° C. FIG. 17A, 1.0067 0.9856 0.0144 0.0039 Anneal at FIG. 17B 1050° C.

[0221] FIG. 22 illustrates a comparison of X-ray diffraction patterns, prepared according to Example #25 and Example #26. The crystal structure of the lithium cobalt oxide materials doped with one or more metal dopants (Example #25 and Example #26) have been investigated by means of X-ray diffraction. Example #25 and Example #26 exhibit a $\text{LiCoO}_{1.2}$ single phase. No second phases, impurity phases, such as $\text{Li}_{0.2}\text{CO}_{0.3}$, $\text{Co}_{0.3}\text{O}_{0.4}$ are observed.

[0222] Details of XRD results, prepared according to Example #25 and Example #26 are shown as Table 22. Comparison of the XRD results, prepared according to Example #25 and Example #26 are shown as Table 23. Based on the XRD results, one observation can be found that the intensity ratio $I(003)/I(104)$ of Example #25 is higher than the intensity ratio $I(003)/I(104)$ of Example #26. Further observation can be found that FWHM (003) and (104) of Example #25 is lower than Example #26. Another observation can be found that c/a of Example #26 is higher than c/a of Example #25. Still another observation can be found that $\Delta 2\theta[(012)-(006)]$, $\Delta 2\theta[(110)-(018)]$ of Example #25 is similar to the data of) Example #26.

TABLE-US-00022 TABLE 22 XRD Results Example#25 Example#26
a [Å] 2.8143 ± 0.002 2.8159 ± 0.0002 (0.07%) (0.007%)
c [Å] 14.044 ± 0.001 14.058 ± 0.001 (0.007%) (0.007%)
c/a 4.990 4.992
hkl 003 104 003 104
2θ[°] 18.955 45.253 18.917 45.225
FWHM[°] 0.09 0.07 0.10 0.09
I (003)/I (104) 3.96 2.27
hkl 006 012 006 012
2θ[°] 38.427 39.090 38.395 39.062
 $\Delta 2\theta[(012)-(016)][°]$ 0.663 0.667
hkl 018 110 018 110
2θ[°] 65.456 66.372 65.431 66.346
 $\Delta 2\theta[(010)-(018)][°]$ 0.916 0.915
r.sub.I 0.51 0.45
K.sub.Fm 23.53 15.37

TABLE-US-00023 TABLE 23 Comparison of the XRD Results
Sample no. Example 25 Example 26
I (003)/I (104) 3.96 2.27
hkl 003 104 003 104
FWHM[°] 0.09 0.07 0.10 0.09
 $\Delta 2\theta[(012)-(006)][°]$ 0.663 0.667
 $\Delta 2\theta[(110)-(118)][°]$ 0.916 0.915
c/a 4.990 4.992
r.sub.I 0.51 0.45
K.sub.Fm 23.53 15.37

Claims

1. A method of producing a lithium cobalt oxide material with one or more metal dopants comprising: forming a mist of a liquid mixture, where the liquid mixture is obtained from adjusting a molar ratio $\text{M.sub.LiSalt}:\text{M.sub.CoSalt}:\text{M.sub.MeSalt}$ of a lithium-containing salt, a cobalt-containing salt, and at least one metal-dopant-containing salt in the liquid mixture to be a ratio of about $x:y:z$, and wherein x from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and $0 < a \leq 0.05$; mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture; drying the gas-liquid mixture to form a gas-solid mixture; separating the gas-solid mixture into one or more solid particles of an oxide material; and annealing the one or more solid particles of the oxide material at an annealing temperature of 400° C. to 1200° C. to obtain crystalized particles of the lithium cobalt oxide material doped with the at least one metal dopant (Me) ($\text{Li}_{0.2}\text{Co}_{0.3}\text{O}_{0.4}\text{Math.doped Me}_{0.2}$), wherein, based on X-ray diffraction (XRD) pattern, the lithium cobalt oxide material doped with the at least one metal dopant exhibits a c/a lattice parameter ratio of not less than 4.990.
2. The method of claim 1, wherein the lithium-containing salt is selected from a group consisting of lithium sulfate ($\text{Li}_{0.2}\text{SO}_{0.4}$), lithium nitrate ($\text{LiNO}_{0.3}$), lithium carbonate ($\text{Li}_{0.2}\text{CO}_{0.3}$), lithium acetate ($\text{LiCH}_{0.2}\text{COO}$), lithium hydroxide (LiOH), lithium formate ($\text{LiCHO}_{0.2}$), lithium chloride (LiCl), and combinations thereof.
3. The method of claim 1, wherein the cobalt-containing salt is selected from a group consisting of cobalt sulfate ($\text{CoSO}_{0.4}$), cobalt nitrate ($\text{Co(NO}_{0.3}\text{)}_{0.2}$), cobalt acetate ($\text{Co(CH}_{0.2}\text{COO)}_{0.2}$), cobalt formate ($\text{Co(CHO}_{0.2}\text{)}_{0.2}$), cobalt chloride ($\text{CoCl}_{0.2}$), and combinations thereof.
4. The method of claim 1, wherein the at least one metal dopant is selected from a group consisting

of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

5. The method of claim 1, wherein the at least one metal-dopant-containing salt is selected from a group consisting of magnesium nitrate $\text{Mg}(\text{NO}_3)_2$, magnesium acetate (MgAc , $\text{Mg}(\text{CH}_3\text{COO})_2$), magnesium chloride (MgCl_2), magnesium sulfate (MgSO_4), magnesium formate ($\text{C}_2\text{H}_2\text{MgO}_4$), aluminum nitrate ($\text{Al}(\text{NO}_3)_3$), aluminum acetate (AlAc , $\text{C}_6\text{H}_9\text{AlO}_6$), aluminum chloride (AlCl_3), aluminum sulfate ($\text{Al}_2(\text{SO}_4)_3$), aluminum formate ($\text{Al}(\text{HCOO})_3$), manganese sulfate (MnSO_4), manganese nitrate ($\text{Mn}(\text{NO}_3)_2$), manganese acetate ($\text{Mn}(\text{CH}_3\text{COO})_2$), manganese formate ($\text{Mn}(\text{CHO}_2)_2$), manganese chloride (MnCl_2), zirconium nitrate ($\text{Zr}(\text{NO}_3)_4$), zirconium acetate ($\text{C}_8\text{H}_{12}\text{O}_8\text{Zr}$), zirconium chloride (ZrCl_4), zirconium sulfate ($\text{Zr}(\text{SO}_4)_2$), zirconium formate ($\text{C}_4\text{H}_4\text{O}_8\text{Zr}$), nickel sulfate (NiSO_4), nickel nitrate ($\text{Ni}(\text{NO}_3)_2$), nickel acetate ($\text{Ni}(\text{CH}_3\text{COO})_2$), nickel formate ($\text{Ni}(\text{CHO}_2)_2$), nickel chloride (NiCl_2), titanyl nitrate ($(\text{TiO}(\text{NO}_3)_2)$), magnesium (Mg)-containing compound, aluminum (Al)-containing compound, titanium (Ti)-containing compound, sodium (Na)-containing compound, potassium (K)-containing compound, scandium (Sc)-containing compound, niobium (Nb)-containing compound, neodymium (Nd)-containing compound, lanthanum (La)-containing compound, cerium (Ce)-containing compound, silicon (Si)-containing compound, rubidium (Rb)-containing compound, vanadium (V)-containing compound, cesium (Cs)-containing compound, chromium (Cr)-containing compound, copper (Cu)-containing compound, magnesium (Mg)-containing compound, manganese (Mn)-containing compound, zirconium (Zr)-containing compound, zinc (Zn)-containing compound, tin (Sn)-containing compound, gallium (Ga)-containing compound, barium (Ba)-containing compound, actinium (Ac)-containing compound, calcium (Ca)-containing compound, iron (Fe)-containing compound, boron (B)-containing compound, germanium (Ge)-containing compound, arsenic (As)-containing compound, hafnium (Hf)-containing compound, Molybdenum (Mo)-containing compound, tungsten (W)-containing compound, rhenium (Re)-containing compound, ruthenium (Ru)-containing compound, rhodium (Rh)-containing compound, platinum (Pt)-containing compound, silver (Ag)-containing compound, osmium (Os)-containing compound, iridium (Ir)-containing compound, gold (Au)-containing compound.

6. The method of claim 1, wherein the liquid mixture is soluble in a suitable solvent and the suitable solvent is selected from a group consisting of water, alcohol, methanol, isopropyl alcohol, organic solvents, inorganic solvents, organic acids, sulfuric acid (H_2SO_4), citric acid ($\text{C}_6\text{H}_8\text{O}_7$), acetic acids (CH_3COOH), butyric acid ($\text{C}_4\text{H}_8\text{O}_2$), lactic acid ($\text{C}_3\text{H}_6\text{O}_3$), nitric acid (HNO_3), hydrochloric acid (HCl), ethanol, pyridine, ammonia, acetone, and combinations thereof.

7. The method of claim 1, wherein the one or more solid particles of the oxide material are annealed in the presence of a second gas flow that is heated to 550° C. or higher and the second gas flow is delivered into a reaction chamber to maintain the annealing temperature inside the reaction chamber.

8. The method of claim 7, wherein the one or more solid particles of the oxide material are annealed in the presence of the second gas flow inside a reaction chamber and the annealing temperature inside the reaction chamber is maintained via a heating element coupled to the reaction chamber.

9. The method of claim 1, wherein the liquid mixture is dried in the presence of the first gas that is heated to 200° C. or higher inside a drying chamber and the first gas is delivered into the drying chamber to maintain a drying temperature inside the drying chamber.

10. The method of claim 1, wherein the liquid mixture is dried inside a drying chamber and a drying temperature inside the drying chamber is maintained via a heating element coupled to the

drying chamber.

11. The method of claim 1, wherein the at least one metal dopant comprises magnesium (Mg), and wherein a battery cell made from the lithium cobalt oxide materials doped with the magnesium, which is annealed at 1090° C., exhibits a first discharge capacity no less than 175 mAh/g and a coulombic efficiency greater than 97% at 4.5 Voltage.

12. The method of claim 1, wherein the at least one metal dopant comprises zirconium (Zr), such that the battery cell made from the lithium cobalt oxide materials doped with the zirconium and annealed at 1020° C. exhibits a first discharge capacity no less than 194 mAh/g and a coulombic efficiency greater than 97% at 4.5 Voltage.

13. The method of claim 1, wherein the at least one metal dopant comprises aluminum (Al), such that the battery cell made from the lithium cobalt oxide materials doped with the zirconium exhibits a first discharge capacity greater than 191 mAh/g and a coulombic efficiency greater than 97% at 4.5 Voltage.

14. A method of producing a lithium cobalt oxide material with one or more metal dopants comprising: forming a mist of a liquid mixture, where the liquid mixture is obtained from adjusting a molar ratio $M_{\text{sub.LiSalt}}:M_{\text{sub.CoSalt}}:M_{\text{sub.Me1Salt}}:M_{\text{sub.Me2Salt}}:M_{\text{sub.Me3Salt}}:\dots:M_{\text{sub.MeNSalt}}$ of a lithium-containing salt, a cobalt-containing salt, and one or more metal-dopant-containing salts which are soluble in a suitable solvent into a liquid mixture to be a ratio of about 1:1:a:b:c : . . . :n, wherein $N \geq 1$, and each a, b, c, . . . , n is more than 0 and not more than 0.05, wherein each of the at least two metal-dopant-containing salts is selected from a group consisting of a first metal-containing salt, a second metal-containing salt, a third metal-containing salt, . . . an N metal-containing salt and combinations thereof; mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture; drying the gas-liquid mixture to form a gas-solid mixture; separating the gas-solid mixture into one or more solid particles of an oxide material; and annealing the one or more solid particles of the oxide material at an annealing temperature of 400° C. to 1200° C. to obtain crystalized particles of the lithium cobalt oxide material doped with the at least one metal dopant (Me) ($\text{Li}_{\text{sub.x}}\text{Co}_{\text{sub.y}}\text{O}_{\text{sub.z}}$ Math.doped $\text{Me}_{\text{sub.a}}$), wherein, based on X-ray diffraction (XRD) pattern, the lithium cobalt oxide material doped with the at least one metal dopant exhibits a c/a lattice parameter ratio of not less than 4.990.

15. The method of claim 14, wherein the lithium-containing salt is selected from a group consisting of lithium sulfate ($\text{Li}_{\text{sub.2}}\text{SO}_{\text{sub.4}}$), lithium nitrate ($\text{LiNO}_{\text{sub.3}}$), lithium carbonate ($\text{Li}_{\text{sub.2}}\text{CO}_{\text{sub.3}}$), lithium acetate ($\text{LiCH}_{\text{sub.2}}\text{COO}$), lithium hydroxide (LiOH), lithium formate ($\text{LiCHO}_{\text{sub.2}}$), lithium chloride (LiCl), and combinations thereof.

16. The method of claim 14, wherein the cobalt-containing salt is selected from a group consisting of cobalt sulfate ($\text{CoSO}_{\text{sub.4}}$), cobalt nitrate ($\text{Co}(\text{NO}_{\text{sub.3}})_{\text{sub.2}}$), cobalt acetate ($\text{Co}(\text{CH}_{\text{sub.2}}\text{COO})_{\text{sub.2}}$), cobalt formate ($\text{Co}(\text{CHO}_{\text{sub.2}})_{\text{sub.2}}$), cobalt chloride ($\text{CoCl}_{\text{sub.2}}$), and combinations thereof.

17. The method of claim 14, wherein each of the one or more metal dopants is selected from a group consisting of Al, Mg, Mn, Zr, Zn, Nb, La, Ce, Sn, Ga, Ba, Ac, Ca, Sc, Ti, V, Cr, Fe, Cu, B, Ge, As, Hf, Mo, W, Re, Ru, Rh, Pt, Ag, Os, Ir, Au, and combinations thereof.

18. The method of claim 14, wherein each of the one or more metal-dopant-containing salts is selected from a group consisting of magnesium nitrate ($\text{Mg}(\text{NO}_{\text{sub.3}})_{\text{sub.2}}$), magnesium acetate (MgAc , $\text{Mg}(\text{CH}_{\text{sub.3}}\text{COO})_{\text{sub.2}}$), magnesium chloride ($\text{MgCl}_{\text{sub.2}}$), magnesium sulfate ($\text{MgSO}_{\text{sub.4}}$), magnesium formate ($\text{C}_{\text{sub.2}}\text{H}_{\text{sub.2}}\text{MgO}_{\text{sub.4}}$), aluminum nitrate ($\text{Al}(\text{NO}_{\text{sub.3}})_{\text{sub.3}}$), aluminum acetate (AlAc , $\text{C}_{\text{sub.6}}\text{H}_{\text{sub.9}}\text{AlO}_{\text{sub.6}}$), aluminum chloride ($\text{AlCl}_{\text{sub.3}}$), aluminum sulfate ($\text{Al}_{\text{sub.2}}(\text{SO}_{\text{sub.4}})_{\text{sub.3}}$), aluminum formate ($\text{Al}(\text{HCOO})_{\text{sub.3}}$), manganese sulfate ($\text{MnSO}_{\text{sub.4}}$), manganese nitrate ($\text{Mn}(\text{NO}_{\text{sub.3}})_{\text{sub.2}}$), manganese acetate ($\text{Mn}(\text{CH}_{\text{sub.2}}\text{COO})_{\text{sub.2}}$), manganese formate ($\text{Mn}(\text{CHO}_{\text{sub.2}})_{\text{sub.2}}$), manganese chloride ($\text{MnCl}_{\text{sub.2}}$), zirconium nitrate ($\text{Zr}(\text{NO}_{\text{sub.3}})_{\text{sub.4}}$), zirconium acetate ($\text{C}_{\text{sub.8}}\text{H}_{\text{sub.12}}\text{O}_{\text{sub.8}}\text{Zr}$), zirconium chloride ($\text{ZrCl}_{\text{sub.4}}$), zirconium sulfate

(Zr(SO₄).sub.4).sub.2), zirconium formate (C.sub.4H.sub.4O.sub.8Zr), nickel sulfate (NiSO₄.sub.4), nickel nitrate (Ni(NO₃).sub.3).sub.2), nickel acetate (Ni(CH₃COO).sub.2), nickel formate (Ni(CHO).sub.2).sub.2), nickel chloride (NiCl.sub.2), titanyl nitrate ((TiO(NO₃).sub.3).sub.2)), magnesium (Mg)-containing compound, aluminum (Al)-containing compound, titanium (Ti)-containing compound, sodium (Na)-containing compound, potassium (K)-containing compound, scandium (Sc)-containing compound, niobium (Nb)-containing compound, neodymium (Nd)-containing compound, lanthanum (La)-containing compound, cerium (Ce)-containing compound, silicon (Si)-containing compound, rubidium (Rb)-containing compound, vanadium (V)-containing compound, cesium (Cs)-containing compound, chromium (Cr)-containing compound, copper (Cu)-containing compound, magnesium (Mg)-containing compound, manganese (Mn)-containing compound, zirconium (Zr)-containing compound, zinc (Zn)-containing compound, tin (Sn)-containing compound, gallium (Ga)-containing compound, barium (Ba)-containing compound, actinium (Ac)-containing compound, calcium (Ca)-containing compound, iron (Fe)-containing compound, boron (B)-containing compound, germanium (Ge)-containing compound, arsenic (As)-containing compound, hafnium (Hf)-containing compound, Molybdenum (Mo)-containing compound, tungsten (W)-containing compound, rhenium (Re)-containing compound, ruthenium (Ru)-containing compound, rhodium (Rh)-containing compound, platinum (Pt)-containing compound, silver (Ag)-containing compound, osmium (Os)-containing compound, iridium (Ir)-containing compound, gold (Au)-containing compound.

19. The method of claim 14, wherein the adjusting of the molar ratio M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt: . . . M.sub.MeNSalt of the lithium-containing salt, the cobalt-containing salt, and the one or more metal-dopant-containing salts is performed prior to forming the mist of the liquid mixture.

20. The method of claim 14, wherein the adjusting of the molar ratio M.sub.LiSalt:M.sub.CoSalt:M.sub.Me1Salt:M.sub.Me2Salt:M.sub.Me3Salt:M.sub.MeNSalt of the lithium-containing salt, the cobalt-containing salt, and the one or more metal-dopant-containing salts is performed simultaneously in forming the mist of the liquid mixture.

21. The method of claim 14, wherein the suitable solvent is selected from a group consisting of water, alcohol, methanol, isopropyl alcohol, organic solvents, inorganic solvents, organic acids, sulfuric acid (H₂SO₄), citric acid (C₆H₈O₇), acetic acids (CH₃COOH), butyric acid (C₄H₈O₂), lactic acid (C₃H₆O₃), Nitric acid (HNO₃), hydrochloric acid (HCl), ethanol, pyridine, ammonia, acetone, and combinations thereof.

22. The method of claim 14, wherein the at least one metal dopant comprises magnesium (Mg) and aluminum (Al), such that the battery cell made from the lithium cobalt oxide materials doped with the magnesium and the aluminum exhibits a first discharge capacity greater than 190 mAh/g and a coulombic efficiency greater than 96% at 4.5 Voltage.

23. A method of producing a lithium cobalt oxide material with one or more metal dopants comprising: forming a mist of a liquid mixture, where the liquid mixture is obtained from adjusting a molar ratio M.sub.LiSalt:M.sub.CoSalt:M.sub.MeSalt of a lithium-containing salt, a cobalt-containing salt, and at least one metal-dopant-containing salt in a suitable solvent into a liquid mixture to be a ratio of about x:y:z:a, and wherein x from 0.9 to 1.1 ($0.9 \leq x \leq 1.1$), y is from 0.9 to 1.1 ($0.9 \leq y \leq 1.1$), z is from 1.8 to 2.2 ($1.8 \leq z \leq 2.2$), and $0 < a \leq 0.05$; mixing the mist of the liquid mixture with a gas flow to form a gas-liquid mixture; drying the gas-liquid mixture to form one or more solid particles of an oxide material; and annealing the one or more solid particles of the oxide material at an annealing temperature of 400° C. to 1200° C. to obtain crystalized particles of the lithium cobalt oxide material doped with the at least one metal dopant (Me) (Li_xCo_yO_z.Math.doped Me_a), wherein, based on X-ray diffraction (XRD) pattern, the lithium cobalt oxide material doped with the at least one metal dopant exhibits a c/a lattice parameter ratio of not less than 4.990.

