PENGERTIAN TABEL PERIODIK DAN HTML, CSS, JSNYA

- **Code Index.html**

```
<div id="status" style="display:none;position:fixed;background:rgba(0, 0, 0, 0.75);color:white;top:0;left:0;width:100%;height:auto;z-index:5;text-align:center;"></div>
<table>
    <caption id="title">
    TABEL PERIODIK UNSUR-UNSUR KIMIA</caption>
    <link rel="stylesheet" href="style.css">
  <thead>
    <tr class="gol">
      <th class="indiv">Golongan
        <br/>Periodik</th>
      <th class="gol">
        <div class="content">1A
          <br/>Alkali</div>
      </th>
      <th class="gol">
        <div class="content">2A
          <br/>Alkali Tanah</div>
      </th>
      <th class="gol">
        <div class="content">3B</div>
      </th>
      <th class="gol">
        <div class="content">4B</div>
      </th>
      <th class="gol">
        <div class="content">5B</div>
      </th>
      <th class="gol">
        <div class="content">6B</div>
      </th>
      <th class="gol">
        <div class="content">7B</div>
      </th>
      <th class="gol" colspan="3">8B</th>
      <th class="gol">
        <div class="content">1B</div>
      </th>
      <th class="gol">
        <div class="content">2B</div>
      </th>
      <th class="gol">
        <div class="content">3A</div>
      </th>
      <th class="gol">
        <div class="content">4A</div>
      </th>
      <th class="gol">
        <div class="content">5A</div>
      </th>
```

```
        <th class="gol">
          <div class="content">6A</div>
        </th>
        <th class="gol">
          <div class="content">7A
            <br/>Halogen</div>
        </th>
        <th class="gol">
          <div class="content">8A
            <br/>Gas Mulia</div>
        </th>
      </tr>
    </thead>
      <tbody>
      <tr>
        <th class="per">
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        </th>
        <td class="g">
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            <div class="ma">1.008</div>
            <br/>
            <div class="lu">H</div>
            <div class="nu">Hidrogen</div>
          </div>
        </td>
        <td colspan="16"></td>
        <td class="gm">
          <div class="content-tp">
            <div class="na">2</div>
            <div class="ma">4.003</div>
            <br/>
            <div class="lu">He</div>
            <div class="nu">Helium</div>
          </div>
        </td>
      </tr>
      <tr>
        <th class="per">
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        </th>
        <td class="l">
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            <div class="ma">6.939</div>
            <br/>
            <div class="lu">Li</div>
            <div class="nu">Litium</div>
          </div>
        </td>
        <td class="l">
          <div class="content-tp">
```

```
            <div class="na">4</div>
            <div class="ma">9.012</div>
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            <div class="lu">Be</div>
            <div class="nu">Berilium</div>
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        <td colspan="10"></td>
        <td class="m">
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            <div class="ma">10.811</div>
            <br/>
            <div class="lu">B</div>
            <div class="nu">Boron</div>
          </div>
        </td>
        <td class="l">
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            <div class="ma">12.011</div>
            <br/>
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            <div class="nu">Karbon</div>
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            <div class="ma">14.007</div>
            <br/>
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            <div class="nu">Nitrogen</div>
          </div>
        </td>
        <td class="g">
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            <div class="na">8</div>
            <div class="ma">15.999</div>
            <br/>
            <div class="lu">O</div>
            <div class="nu">Oksigen</div>
          </div>
        </td>
        <td class="h">
          <div class="content-tp">
            <div class="na">9</div>
            <div class="ma">18.998</div>
            <br/>
            <div class="lu">F</div>
            <div class="nu">Fluor</div>
          </div>
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```

```
        <td class="gm">
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            <div class="ma">20.18</div>
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            <div class="nu">Neon</div>
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        </td>
      </tr>
      <tr>
        <th class="per">
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        </th>
        <td class="l">
          <div class="content-tp">
            <div class="na">11</div>
            <div class="ma">22.989</div>
            <br/>
            <div class="lu">Na</div>
            <div class="nu">Natrium</div>
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        <td class="l">
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            <div class="ma">24.305</div>
            <br/>
            <div class="lu">Mg</div>
            <div class="nu">Magnesium</div>
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        <td colspan="10"></td>
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            <div class="na">13</div>
            <div class="ma">26.981</div>
            <br/>
            <div class="lu">Al</div>
            <div class="nu">Aluminium</div>
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        </td>
        <td class="m">
          <div class="content-tp">
            <div class="na">14</div>
            <div class="ma">28.086</div>
            <br/>
            <div class="lu">Si</div>
            <div class="nu">Silikon</div>
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        </td>
        <td class="l">
          <div class="content-tp">
```

```
            <div class="na">15</div>
            <div class="ma">30.974</div>
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            <div class="nu">Fosfor</div>
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          <div class="content-tp">
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            <div class="ma">32.066</div>
            <br/>
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            <div class="nu">Belerang</div>
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            <div class="na">17</div>
            <div class="ma">35.453</div>
            <br/>
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            <div class="nu">Klor</div>
          </div>
        </td>
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            <div class="ma">39.948</div>
            <br/>
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            <div class="nu">Argon</div>
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        </th>
        <td class="l">
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            <div class="na">19</div>
            <div class="ma">39.098</div>
            <br/>
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            <div class="nu">Kalium</div>
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        </td>
        <td class="l">
          <div class="content-tp">
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            <div class="ma">40.076</div>
            <br/>
```

```
            <div class="lu">Ca</div>
            <div class="nu">Kalsium</div>
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        <td class="l">
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            <div class="ma">44.956</div>
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            <div class="nu">Skandium</div>
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        <td class="l">
          <div class="content-tp">
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            <div class="ma">47.88</div>
            <br/>
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            <div class="nu">Titanium</div>
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        </td>
        <td class="l">
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            <div class="ma">50.942</div>
            <br/>
            <div class="lu">V</div>
            <div class="nu">Vanadium</div>
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        </td>
        <td class="l">
          <div class="content-tp">
            <div class="na">24</div>
            <div class="ma">51.996</div>
            <br/>
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            <div class="nu">Kromium</div>
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        </td>
        <td class="l">
          <div class="content-tp">
            <div class="na">25</div>
            <div class="ma">54.938</div>
            <br/>
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            <div class="nu">Mangan</div>
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          <div class="content-tp">
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            <div class="ma">55.847</div>
```

```
            <br/>
            <div class="lu">Fe</div>
            <div class="nu">Besi</div>
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        </td>
        <td class="l">
          <div class="content-tp">
            <div class="na">27</div>
            <div class="ma">58.933</div>
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            <div class="nu">Kobalt</div>
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        <td class="l">
          <div class="content-tp">
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            <div class="ma">58.69</div>
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            <div class="nu">Nikel</div>
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        <td class="l">
          <div class="content-tp">
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            <div class="ma">63.546</div>
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            <div class="lu">Cu</div>
            <div class="nu">Tembaga</div>
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        </td>
        <td class="l">
          <div class="content-tp">
            <div class="na">30</div>
            <div class="ma">65.39</div>
            <br/>
            <div class="lu">Zn</div>
            <div class="nu">Seng</div>
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        <td class="c">
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            <div class="ma">69.723</div>
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            <div class="lu">Ga</div>
            <div class="nu">Galium</div>
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        <td class="m">
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```

```
            <div class="ma">72.61</div>
            <br/>
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            <div class="nu">Germanium</div>
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          <div class="content-tp">
            <div class="na">33</div>
            <div class="ma">74.922</div>
            <br/>
            <div class="lu">As</div>
            <div class="nu">Arsen</div>
          </div>
        </td>
        <td class="l">
          <div class="content-tp">
            <div class="na">34</div>
            <div class="ma">78.96</div>
            <br/>
            <div class="lu">Se</div>
            <div class="nu">Selenium</div>
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        </td>
        <td class="h">
          <div class="content-tp">
            <div class="na">35</div>
            <div class="ma">79.904</div>
            <br/>
            <div class="lu">Br</div>
            <div class="nu">Bromin</div>
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        </td>
        <td class="gm">
          <div class="content-tp">
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            <div class="ma">83.8</div>
            <br/>
            <div class="lu">Kr</div>
            <div class="nu">Kripton</div>
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      <tr>
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            <div class="ma">85.4678</div>
            <br/>
            <div class="lu">Rb</div>
            <div class="nu">Rubidium</div>
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```

```
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            <div class="nu">Stronsium</div>
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        </td>
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            <div class="ma">88.906</div>
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            <div class="lu">Y</div>
            <div class="nu">Itrium</div>
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            <div class="ma">91.224</div>
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            <div class="nu">Zikronium</div>
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            <div class="ma">92.960</div>
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            <div class="nu">Niobium</div>
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        </td>
        <td class="l">
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            <div class="ma">95.94</div>
            <br/>
            <div class="lu">Mo</div>
            <div class="nu">Moilbdenum</div>
          </div>
        </td>
        <td class="b">
          <div class="content-tp">
            <div class="na">43</div>
            <div class="ma">98 (0)</div>
            <br/>
            <div class="lu">Tc</div>
            <div class="nu">Teknesium</div>
```

```
          </div>
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            <div class="ma">102.905</div>
            <br/>
            <div class="lu">Rd</div>
            <div class="nu">Rodium</div>
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        <td class="l">
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            <div class="ma">106.42</div>
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            <div class="nu">Rubidium</div>
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        <td class="l">
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            <div class="na">47</div>
            <div class="ma">107.868</div>
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            <div class="nu">Perak</div>
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            <div class="ma">112.41</div>
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            <div class="nu">Kadmium</div>
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        </td>
        <td class="l">
          <div class="content-tp">
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            <div class="ma">114.82</div>
            <br/>
            <div class="lu">In</div>
```

```
            <div class="nu">Indium</div>
          </div>
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            <br/>
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            <div class="nu">Timah</div>
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            <div class="ma">121.76</div>
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            <div class="nu">Antimon</div>
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          <div class="content-tp">
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            <br/>
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            <div class="nu">Telurium</div>
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          <div class="content-tp">
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            <div class="ma">126.904</div>
            <br/>
            <div class="lu">I</div>
            <div class="nu">Yodium</div>
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            <div class="ma">131.29</div>
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            <div class="lu">Xe</div>
            <div class="nu">Xenon</div>
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        </td>
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      <tr>
        <th class="per">
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```

```
        <td class="c">
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            <div class="nu">Sesium</div>
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        </td>
        <td class="l">
          <div class="content-tp">
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            <div class="ma">137.327</div>
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        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">57-71</div>
            <div class="ma">138-174</div>
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            <div class="nu">Rangkaian Lantanida</div>
            <br/>
            <div class="nu-desc">Nomor 57 sampai 71</div>
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            <div class="nu">Tantalum</div>
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        <td class="l">
          <div class="content-tp">
            <div class="na">74</div>
            <div class="ma">137.327</div>
            <br/>
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            <div class="nu">Barium</div>
```

```
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            <div class="nu">Renium</div>
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        <td class="l">
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            <div class="nu">Irdium</div>
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        <td class="l">
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            <div class="nu">Emas</div>
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            <div class="ma">200.59</div>
            <br/>
            <div class="lu">Hg</div>
```

```
            <div class="nu">Air Raksa</div>
          </div>
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            <div class="ma">204.37</div>
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            <div class="nu">Talium</div>
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            <div class="nu">Timbal</div>
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            <div class="ma">208.98</div>
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            <div class="nu">Bismut</div>
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        <td class="m">
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    Radioaktif">
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            <div class="ma">(209)</div>
            <br/>
            <div class="lu">Po</div>
            <div class="nu">Barium</div>
          </div>
        </td>
        <td class="h">
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
    Radioaktif">
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            <div class="ma">(210)</div>
            <br/>
            <div class="lu">At</div>
            <div class="nu">Astatin</div>
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        <td class="gm">
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```

```
            <div class="ma">(222)</div>
            <br/>
            <div class="lu">Rn</div>
            <div class="nu">Radon</div>
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        <td class="c">
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Radioaktif">
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            <div class="ma">(223.02)</div>
            <br/>
            <div class="lu">Fr</div>
            <div class="nu">Fransium</div>
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        </td>
        <td class="l">
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
Radioaktif">
            <div class="na">88</div>
            <div class="ma">(226)</div>
            <br/>
            <div class="lu">Ra</div>
            <div class="nu">Radium</div>
          </div>
        </td>
        <td class="akt">
          <div class="content-tp" title="Semua elemen Aktinida termasuk dalam
kategori Radioaktif">
            <div class="na">89-103</div>
            <div class="ma">227-262</div>
            <br/>
            <div class="nu">Rangkaian Aktinida</div>
            <br/>
            <div class="nu-desc">Nomor 89 sampai 103</div>
          </div>
        </td>
        <td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
            <div class="na">104</div>
            <div class="ma">[267]</div>
            <br/>
            <div class="lu">Rf</div>
            <div class="nu">Rutherfordium</div>
          </div>
        </td>
        <td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
```

```
            <div class="na">105</div>
            <div class="ma">[268]</div>
            <br/>
            <div class="lu">Db</div>
            <div class="nu">Dubnium</div>
          </div>
        </td>
        <td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
            <div class="na">106</div>
            <div class="ma">[269]</div>
            <br/>
            <div class="lu">Sg</div>
            <div class="nu">Seaborgium</div>
          </div>
        </td>
        <td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
            <div class="na">107</div>
            <div class="ma">[270]</div>
            <br/>
            <div class="lu">Bh</div>
            <div class="nu">Bohrium</div>
          </div>
        </td>
        <td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
            <div class="na">108</div>
            <div class="ma">[269]</div>
            <br/>
            <div class="lu">Hs</div>
            <div class="nu">Hassium</div>
          </div>
        </td>
        <td class="undef" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
            <div class="na">109</div>
            <div class="ma">[278]</div>
            <br/>
            <div class="lu">Mt</div>
            <div class="nu">Meitnerium</div>
          </div>
        </td>
        <td class="undef" title="Elemen ini termasuk dalam kategori Radioaktif">
          <div class="content-tp">
            <div class="na">110</div>
            <div class="ma">[281]</div>
            <br/>
            <div class="lu" title="Kode lain: Uun">Ds</div>
            <div class="nu" title="Nama lain: Ununnilium">Darmstadtium</div>
          </div>
        </td>
        <td class="undef">
```

```
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
    Radioaktif">
            <div class="na">111</div>
            <div class="ma">[281]</div>
            <br/>
            <div class="lu" title="Kode lain: Uuu">Rg</div>
            <div class="nu" title="Nama lain: Unununium">Roentgenium</div>
          </div>
        </td>
        <td class="l">
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
    Radioaktif">
            <div class="na">112</div>
            <div class="ma">[285]</div>
            <br/>
            <div class="lu" title="Kode lain: Uub">Cn</div>
            <div class="nu" title="Nama lain: Ununbium">Kopernesium</div>
          </div>
        </td>
        <td class="undef">
          <div class="content-tp">
            <div class="na">113</div>
            <div class="ma">[284]</div>
            <br/>
            <div class="lu" title="Kode lain: -">Uut</div>
            <div class="nu" title="Nama lain: -">Ununtrium</div>
          </div>
        </td>
        <td class="undef">
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
    Radioaktif">
            <div class="na">114</div>
            <div class="ma">[289]</div>
            <br/>
            <div class="lu" title="Kode lain: Uuq">Fl</div>
            <div class="nu" title="Nama lain: Ununquadium">Flerovium</div>
          </div>
        </td>
        <td class="undef">
          <div class="content-tp">
            <div class="na">115</div>
            <div class="ma">[288]</div>
            <br/>
            <div class="lu" title="Kode lain: -">Uup</div>
            <div class="nu" title="Nama lain: -">Ununpentium</div>
          </div>
        </td>
        <td class="undef">
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
    Radioaktif">
            <div class="na">116</div>
            <div class="ma">[293]</div>
            <br/>
```

```
            <div class="lu" title="Kode lain: Uuh">Lv</div>
            <div class="nu" title="Nama lain: Ununheksium">Livermorium</div>
          </div>
        </td>
        <td class="undef">
          <div class="content-tp">
            <div class="na">117</div>
            <div class="ma">[294]</div>
            <br/>
            <div class="lu" title="Kode lain: -">Uus</div>
            <div class="nu" title="Nama lain: -">Ununseptium</div>
          </div>
        </td>
        <td class="undef">
          <div class="content-tp" title="Elemen ini termasuk dalam kategori
Radioaktif">
            <div class="na">118</div>
            <div class="ma">[294]</div>
            <br/>
            <div class="lu" title="Kode lain: -">Uuo</div>
            <div class="nu" title="Nama lain: -">Ununoktium</div>
          </div>
        </td>
      </tr>
      <tr>
        <th></th>
        <td colspan="18">
          <div class="content-tp"></div>
        </td>
      </tr>
      <tr>
        <th></th>
        <td></td>
        <td class="lan">
          <div class="content-tp">
            <div class="na"> </div>
            <div class="ma"> </div>
            <br/>
            <div class="nu rlra">Rangkaian Lantanida</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">57</div>
            <div class="ma">138.906</div>
            <br/>
            <div class="lu">La</div>
            <div class="nu">Lantanum</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">58</div>
```

```
            <div class="ma">140.115</div>
            <br/>
            <div class="lu">Ce</div>
            <div class="nu">Serium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">59</div>
            <div class="ma">140.908</div>
            <br/>
            <div class="lu">Pr</div>
            <div class="nu">Praseodimium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">60</div>
            <div class="ma">144.24</div>
            <br/>
            <div class="lu">Nd</div>
            <div class="nu">Neodimium</div>
          </div>
        </td>
        <td class="lan-b">
          <div class="content-tp">
            <div class="na">61</div>
            <div class="ma">(145)</div>
            <br/>
            <div class="lu">Pm</div>
            <div class="nu">Prometium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">62</div>
            <div class="ma">150.36</div>
            <br/>
            <div class="lu">Sm</div>
            <div class="nu">Samarium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">63</div>
            <div class="ma">151.96</div>
            <br/>
            <div class="lu">Eu</div>
            <div class="nu">Europium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
```

```
            <div class="na">64</div>
            <div class="ma">157.25</div>
            <br/>
            <div class="lu">Gd</div>
            <div class="nu">Gadolinium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">65</div>
            <div class="ma">158.923</div>
            <br/>
            <div class="lu">Tb</div>
            <div class="nu">Terbium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">66</div>
            <div class="ma">162.5</div>
            <br/>
            <div class="lu">Dy</div>
            <div class="nu">Disprosium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">67</div>
            <div class="ma">164.93</div>
            <br/>
            <div class="lu">Ho</div>
            <div class="nu">Holmium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">68</div>
            <div class="ma">167.26</div>
            <br/>
            <div class="lu">Er</div>
            <div class="nu">Erbium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">69</div>
            <div class="ma">168.93</div>
            <br/>
            <div class="lu">Tm</div>
            <div class="nu">Tulium</div>
          </div>
        </td>
        <td class="lan">
```

```
          <div class="content-tp">
            <div class="na">70</div>
            <div class="ma">173.04</div>
            <br/>
            <div class="lu">Yb</div>
            <div class="nu">Iterbium</div>
          </div>
        </td>
        <td class="lan">
          <div class="content-tp">
            <div class="na">71</div>
            <div class="ma">174.967</div>
            <br/>
            <div class="lu">Lu</div>
            <div class="nu">Lutelium</div>
          </div>
        </td>
      </tr>
      <tr>
        <th></th>
        <td></td>
        <td class="akt">
          <div class="content-tp">
            <div class="na"> </div>
            <div class="ma"> </div>
            <br/>
            <div class="nu rlra">Rangkaian Aktinida</div>
          </div>
        </td>
        <td class="akt">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
    Radioaktif">
            <div class="na">89</div>
            <div class="ma">(227)</div>
            <br/>
            <div class="lu">Ac</div>
            <div class="nu">Aktinium</div>
          </div>
        </td>
        <td class="akt">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
    Radioaktif">
            <div class="na">90</div>
            <div class="ma">232.038</div>
            <br/>
            <div class="lu">Th</div>
            <div class="nu">Torium</div>
          </div>
        </td>
        <td class="akt">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
    Radioaktif">
            <div class="na">91</div>
```

```
            <div class="ma">231.035</div>
            <br/>
            <div class="lu">Pa</div>
            <div class="nu">Protaktinium</div>
          </div>
        </td>
        <td class="akt">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
Radioaktif">
            <div class="na">92</div>
            <div class="ma">238.029</div>
            <br/>
            <div class="lu">U</div>
            <div class="nu">Uranium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
Radioaktif">
            <div class="na">93</div>
            <div class="ma">(237)</div>
            <br/>
            <div class="lu">Np</div>
            <div class="nu">Neptunium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
Radioaktif">
            <div class="na">94</div>
            <div class="ma">(244)</div>
            <br/>
            <div class="lu">Pu</div>
            <div class="nu">Plutonium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
Radioaktif">
            <div class="na">95</div>
            <div class="ma">(243)</div>
            <br/>
            <div class="lu">Am</div>
            <div class="nu">Amerisium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori
Radioaktif">
            <div class="na">96</div>
            <div class="ma">(247)</div>
            <br/>
            <div class="lu">Cm</div>
```

```
            <div class="nu">Kurium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">97</div>
            <div class="ma">(247)</div>
            <br/>
            <div class="lu">Bk</div>
            <div class="nu">Berkelium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">98</div>
            <div class="ma">(251)</div>
            <br/>
            <div class="lu">Cf</div>
            <div class="nu">Kalifornium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">99</div>
            <div class="ma">(252)</div>
            <br/>
            <div class="lu">Es</div>
            <div class="nu">Einsteinium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">100</div>
            <div class="ma">(257)</div>
            <br/>
            <div class="lu">Fm</div>
            <div class="nu">Fermium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">101</div>
            <div class="ma">(258)</div>
            <br/>
            <div class="lu">Md</div>
            <div class="nu">Mendelevium</div>
          </div>
        </td>
```

```
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">102</div>
            <div class="ma">(259)</div>
            <br/>
            <div class="lu">No</div>
            <div class="nu">Nobelium</div>
          </div>
        </td>
        <td class="akt-b">
          <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
            <div class="na">103</div>
            <div class="ma">(262)</div>
            <br/>
            <div class="lu">Lr</div>
            <div class="nu">Lawrensium</div>
          </div>
        </td>
      </tr>
    </tbody>
</table>
```

- **APP.JS**

```
function showstatus(text, color) {
    $('#status').stop().slideUp().css('color', color).html(text).slideDown().delay(2000).slideUp()
}

window.onbeforeunload = function () {
    if (navigator.onLine) {} else {
        showstatus('<marquee>No network connection. Please do not refresh until this message is disappear</marquee>', 'red')
    }
}

window.oncontextmenu = function () {
    showstatus('Right click is disabled', '#F44336');
    return false
}

shortcut.add('CTRL+U', function () {
    showstatus('View source is disabled!', '#F44336')
}),
shortcut.add('CTRL+Shift+I', function () {
    showstatus('Inspect Element is disabled!', '#F44336')
}),
shortcut.add('CTRL+Shift+J', function () {
    showstatus('JS Console is disabled!', '#F44336')
}),
```

```
shortcut.add('CTRL+Shift+C', function () {
    showstatus('Inspect Element is disabled!', '#F44336')
}),
shortcut.add('F12', function () {
    showstatus('JS Console is disabled!', '#F44336')
}),
shortcut.add('Meta+Alt+U', function () {
    showstatus('View source is disabled!', '#F44336')
}),
shortcut.add('Meta+Alt+I', function () {
    showstatus('Inspect Element is disabled!', '#F44336')
}),
shortcut.add('Meta+Alt+J', function () {
    showstatus('JS Console is disabled!', '#F44336')
}),
shortcut.add('Meta+Shift+C', function () {
    showstatus('Inspect Element is disabled!', '#F44336')
}),
shortcut.add('Meta+P', function () {
    showstatus('Preparing to print...', 'white');
    setTimeout(function () {
        window.print()
    }, 3000)
}),
shortcut.add('Ctrl+P', function () {
    showstatus('Preparing to print...', 'white');
    setTimeout(function () {
        window.print()
    }, 3000)
});
```

- **STYLE.CSS**

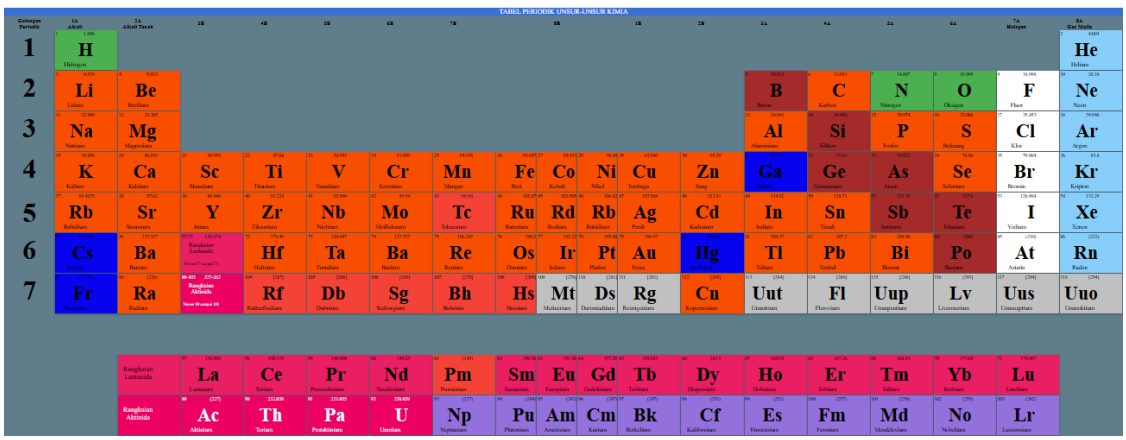
```
body {
    font-size: 11pt;
    line-height: initial;
  }

  table {
    width: 96%;
    background: #607D8B;
    font-family: roboto;
    border-style: none;
    border-collapse: separate;
    border-spacing: 2px;
  }

  th {
    width: 400px;
    max-width: 400px;
    font-family: roboto;
    text-align: center;
  }
```

```
    abbr {
      text-decoration: underline;
    }
    /* Layouting */

    #title {
      background-color: #3870d8;
      color: white;
      text-align: center;
    }

    .indiv {
      font-size: 0.9em;
      font-weight: bold;
    }

    .content {
      width: 90px;
      max-width: 90px;
    }

    .content-per {
      height: 90px;
      max-height: 90px;
    }

    .content-tp {
      width: 90px;
      max-width: 90px;
      height: 90px;
      max-height: 90px;
    }

    .gol {
      font-size: 0.9em;
      font-weight: bold;
    }

    .per {
      font-size: 4em;
      font-weight: bold;
    }
    /* For periodic Elements Inside */

    .na {
      font-size: 0.7em;
      float: left;
    }

    .ma {
      font-size: 0.7em;
      float: right;
```

```
    }

    .lu {
      font-size: 2.75em;
      font-weight: bold;
      text-align: right;
      width: 97%;
      margin-right: 5px;
    }

    .sp {
      font-size: 0.9em;
      float: left;
    }

    .nu {
      font-size: 10.1pt;
      text-align: center;
    }

    .rlra {
      font-size: 12pt;
      margin-top: 5px;
    }
    td
    {
      box-shadow:0 1.5px 4px rgba(0, 0, 0, 0.24), 0 1.5px 6px rgba(0, 0, 0, 0.12)
    }
    td[colspan], td:empty
    {
      box-shadow:none;
    }
    .nu-desc {
      font-size: 7pt;
      text-align: center;
    }

    .kotak {
      width: 85px;
      height: 85px;
    }
    /* CSS for Element Category Starts */

    .l {
      background: #f74e00;
    }

    .c {
      background: #0703ee;
    }

    .g {
      background: #4CAF50;
```

```
    }

    .b {
      background: #F44336;
    }

    .h {
      background: #ffffff;
    }

    .lan {
      background: #E91E63;
    }

    .lan-b {
      background: #F44336;
    }

    .akt {
      background: #ec0062;
      color: #fff;
    }

    .akt-b {
      background: mediumpurple;
    }

    .m {
      background: brown;
    }

    .gm {
      background: LightskyBlue;
    }

    .undef {
      background: silver;
    }
```

- **GAMBAR :**



- **PENJELSAN DI DALAM TABEL PERIODIK BENTUK HTML :**

```
<!DOCTYPE html>
<html lang="id">
<head>
  <meta charset="UTF-8">
  <title>Tabel Periodik dan Fungsinya</title>
  <style>
    body {
      font-family: Arial, sans-serif;
      background-color: #eef3f7;
      padding: 20px;
    }
    h1 {
      text-align: center;
      color: #2c3e50;
    }
    table {
      width: 100%;
      border-collapse: collapse;
      margin-bottom: 30px;
    }
    th, td {
      border: 1px solid #ccc;
      padding: 10px;
      text-align: left;
    }
    th {
      background-color: #2c3e50;
      color: white;
    }
    .color-box {
      display: inline-block;
      width: 20px;
      height: 20px;
      vertical-align: middle;
      margin-right: 10px;
      border: 1px solid #000;
    }
  </style>
</head>
<body>

<h1>Tabel Periodik Unsur dan Fungsinya</h1>

<table>
  <thead>
    <tr>
```

```
    <th>Warna</th>
    <th>Kelompok Unsur</th>
    <th>Contoh Unsur</th>
    <th>Fungsi Umum</th>
  </tr>
</thead>
<tbody>
  <tr>
    <td><span class="color-box" style="background-color: #ff9900;"></span></td>
    <td>Logam Transisi</td>
    <td>Fe, Cu, Zn, Au</td>
    <td>Struktur bangunan, konduktor listrik, perhiasan</td>
  </tr>
  <tr>
    <td><span class="color-box" style="background-color: #ff3300;"></span></td>
    <td>Logam Pasca Transisi</td>
    <td>Al, Sn, Pb</td>
    <td>Kabel listrik, pelapis kaleng, baterai</td>
  </tr>
  <tr>
    <td><span class="color-box" style="background-color: #ff3399;"></span></td>
    <td>Lantanida</td>
    <td>La, Ce, Nd</td>
    <td>Magnet, lampu neon, laser</td>
  </tr>
  <tr>
    <td><span class="color-box" style="background-color: #cc0066;"></span></td>
    <td>Aktinida</td>
    <td>U, Pu, Th</td>
    <td>Reaktor nuklir, senjata nuklir</td>
  </tr>
  <tr>
    <td><span class="color-box" style="background-color: #00ccff;"></span></td>
    <td>Gas Mulia</td>
    <td>He, Ne, Ar</td>
    <td>Balon udara, lampu neon, atmosfer pelindung</td>
  </tr>
  <tr>
    <td><span class="color-box" style="background-color: #66ff66;"></span></td>
    <td>Non-Logam</td>
    <td>H, C, N, O</td>
    <td>Pernapasan, DNA, pupuk, air</td>
  </tr>
  <tr>
    <td><span class="color-box" style="background-color: #cc0000;"></span></td>
    <td>Metaloid</td>
    <td>B, Si, As</td>
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    <td>Semikonduktor, kaca, pestisida</td>
   </tr>
   <tr>
    <td><span class="color-box" style="background-color: #3399ff;"></span></td>
    <td>Logam Alkali</td>
    <td>Li, Na, K</td>
    <td>Baterai, garam, pupuk</td>
   </tr>
   <tr>
    <td><span class="color-box" style="background-color: #ff9966;"></span></td>
    <td>Logam Alkali Tanah</td>
    <td>Mg, Ca, Ba</td>
    <td>Obat, tulang, kembang api</td>
   </tr>
   <tr>
    <td><span class="color-box" style="background-color: #9966cc;"></span></td>
    <td>Halogen</td>
    <td>F, Cl, I</td>
    <td>Disinfektan, garam, antiseptik</td>
   </tr>
   <tr>
    <td><span class="color-box" style="background-color: #cccccc;"></span></td>
    <td>Unsur Sintetik</td>
    <td>Uut, Uus, Uup</td>
    <td>Penelitian ilmiah, tidak stabil</td>
   </tr>
  </tbody>
</table>

<p>Warna-warna di atas mencerminkan kategori unsur sesuai dengan warna pada gambar tabel periodik yang kamu unggah. Setiap warna memiliki fungsi dan kegunaan berbeda dalam kehidupan sehari-hari maupun industri.</p>

</body>
</html>
```

- **PENJELASAN ARTI UNSUR-UNSUR KIMIA PADA TABEL PERIODIK**

🔶 1. Logam Alkali (Warna Biru Cerah – Contoh: Li, Na, K, Rb, Cs, Fr)

- Litium (Li): Digunakan dalam baterai isi ulang (lithium-ion).
- Natrium (Na): Komponen utama garam dapur (NaCl), penting untuk keseimbangan cairan tubuh.
- Kalium (K): Mineral esensial untuk fungsi otot dan jantung.
- Rubidium (Rb), Sesium (Cs), Fransium (Fr): Lebih reaktif, jarang digunakan secara luas, digunakan dalam riset.

🟧 2. Logam Alkali Tanah (Warna Oranye Tua – Contoh: Be, Mg, Ca, Sr, Ba, Ra)

- Magnesium (Mg): Digunakan dalam paduan logam ringan, dan suplemen.
- Kalsium (Ca): Pembentuk tulang dan gigi, penting dalam kontraksi otot.

- Barium (Ba): Digunakan dalam foto rontgen perut (barium meal).
- Radium (Ra): Radioaktif, dulu digunakan dalam cat luminesen (tidak lagi karena bahaya radiasi).

🔴 3. Logam Transisi (Warna Oranye – Contoh: Fe, Cu, Zn, Au, Ag, Ni, Co, Ti, Cr)
- Besi (Fe): Bahan utama baja, penting untuk hemoglobin dalam darah.
- Tembaga (Cu): Konduktor listrik, digunakan dalam kabel.
- Zink (Zn): Melindungi besi dari korosi (galvanisasi), penting untuk sistem imun.
- Perak (Ag): Perhiasan, fotografi, alat medis karena sifat antimikroba.
- Emas (Au): Perhiasan, investasi, digunakan dalam elektronik.

🟤 4. Logam Pasca Transisi (Warna Merah Gelap – Contoh: Al, Sn, Pb, Bi, In)
- Aluminium (Al): Ringan dan tahan karat, digunakan dalam kemasan, pesawat.
- Timah (Sn): Pelapis kaleng makanan, solder elektronik.
- Timbal (Pb): Dulu digunakan dalam cat dan bensin (sekarang dilarang), masih digunakan dalam aki.
- Bismut (Bi): Obat maag (seperti Pepto-Bismol).

🟣 5. Metaloid (Warna Ungu Merah – Contoh: B, Si, As, Sb, Te)
- Boron (B): Digunakan dalam serat borosilikat (Pyrex).
- Silikon (Si): Bahan utama chip komputer dan panel surya.
- Arsenik (As): Sangat beracun, pernah digunakan dalam pestisida.

🟩 6. Non-Logam (Warna Hijau – Contoh: H, C, N, O, P, S, Se)
- Hidrogen (H): Unsur paling ringan, bahan bakar potensial.
- Karbon (C): Dasar semua senyawa organik (makhluk hidup).
- Nitrogen (N): 78% udara, bahan baku pupuk.
- Oksigen (O): Diperlukan untuk pernapasan.
- Fosfor (P): Pupuk, DNA, tulang.
- Sulfur (S): Dalam obat dan bahan peledak (mesiu hitam).
- Selenium (Se): Mikronutrien, digunakan dalam elektronik.

🟪 7. Halogen (Warna Ungu Muda – Contoh: F, Cl, Br, I, At)
- Fluorin (F): Dalam pasta gigi untuk mencegah gigi berlubang.
- Klorin (Cl): Disinfektan dalam air kolam renang.
- Bromin (Br): Digunakan dalam bahan pemadam api.
- Iodin (I): Penting untuk kelenjar tiroid.

🟦 8. Gas Mulia (Warna Biru Muda – Contoh: He, Ne, Ar, Kr, Xe, Rn)
- Helium (He): Balon, tidak mudah terbakar.
- Neon (Ne): Lampu reklame.
- Argon (Ar): Gas inert dalam lampu bohlam dan pengelasan.
- Radon (Rn): Gas radioaktif alami, bisa berbahaya jika terakumulasi.

🟥 9. Lantanida (Baris Atas Merah Muda – Contoh: La – Lu)
- Neodimium (Nd): Magnet super kuat.
- Europium (Eu): Digunakan dalam layar TV dan lampu fluoresen.

🟥 10. Aktinida (Baris Bawah Merah Tua – Contoh: Ac – Lr)

- Uranium (U): Bahan bakar reaktor nuklir.
- Plutonium (Pu): Digunakan dalam senjata nuklir dan reaktor.

⚙️ 11. Unsur Sintetik (Warna Abu-Abu – Contoh: Uut, Uup, Uus, Ds, Rg, Og, dll.)

- Unsur yang tidak ditemukan di alam dan hanya dibuat di laboratorium.
- Umumnya bersifat radioaktif dan tidak stabil.
- Digunakan dalam riset ilmiah dan fisika partikel.