Molecular mass of SO_2 = 32 + (2×16) = 64amu

QUESTION

Write the molecular formula and calculate the molecular mass of the following compounds:

(a) Sulphur trioxide
(b) Iron II Sulphide
(c) Ammonia

ans (a) Sulphur trioxide (SO_3)

Molecular mass = 32 + (3×16) = 80 amu

(b) Iron (II) Sulphide (FeS)

Molecular mass = 56 + 32 = 88amu

(c) Ammonia (NH_3)

Molecular mass = 14 + (3×1) = 17amu