



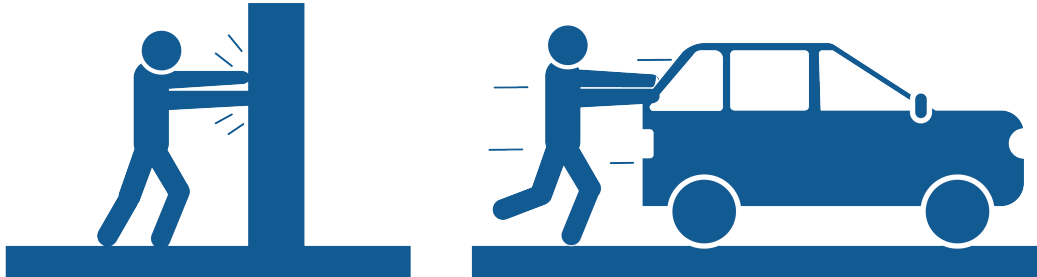
## CHAPTER 1

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# Work and Energy

In this chapter, we are going to talk about how engineers define work and energy. It frequently takes force to get work done. Let's start with thinking about the relationship between force and energy. As we learned earlier, Force is measured in newtons, and one newton is equal to the force necessary to accelerate one kilogram at a rate of  $1\text{m/s}^2$ .

When you lean on a wall, you are exerting a force on the wall, but you aren't doing any work. On the other hand, if you push a car for a mile, you are clearly doing work. Work, to an engineer, is the force you apply to something, as well as the distance that it moves, in the direction of the applied force. We measure work in *joules*. A joule is one newton of force over one meter.



For example, if you push a car uphill with a force of 10 newtons for 12 meters, you have done 120 joules of work.

Work is how energy is transferred from one thing to another. When you push the car, you also burn sugars(energy of the body) in your blood. That energy is then transferred to the car: after it has been pushed uphill.

Thus, we measure the energy something consumes or generates in units of work: joules, kilowatt-hours, horsepower-hours, foot-pounds, BTUs( British Thermal Unit), and calories.

Let's go over a few different forms that energy can take.

Watch Khan Academy's **Changes in energy** at <https://www.khanacademy.org/science/ms-physics/x1baed5db7c1bb50b:energy/x1baed5db7c1bb50b:changes-in-energy/a/changes-in-energy>

### 1.1 Forms of Energy

In this section we are going to learn about several different types of energy:

- Heat
- Chemical Energy
- Kinetic Energy
- Gravitational Potential Energy

### 1.1.1 Heat

When you heat something, you are transferring energy to it. The BTU is a common unit for heat: One BTU is the amount of heat required to raise the temperature of one pound of water, by one degree. One BTU is about 1,055 joules. In fact, when you buy and sell natural gas as fuel, it is priced by the BTU.

### 1.1.2 Electricity

Electricity is the movement of electrons. When you push electrons through a space that resists their passage (like a light bulb), energy is transferred from the power source ( a battery) into the source of the resistance.

Let's say your lightbulb consumes 60 watts of electricity, and you leave it on for 24 hours. We would say that you have consumed 1.44 kilowatt hours or 3,600,000 joules.

Watch Khan Academy's **Introduction to charge** at <https://www.khanacademy.org/science/in-in-class10th-physics/in-in-electricity/in-in-electric-current-circuit/v/intro-to-charge>

### 1.1.3 Chemical Energy

As mentioned early, some chemical reactions consume energy and some produce energy. Thus, energy can be stored in the structure of a molecule. When a plant uses photosynthesis to rearrange water and carbon dioxide into a sugar molecule, it converts the energy in the sunlight( solar energy) into chemical energy. Remember photosynthesis is a process that releases energy. Therefore, the sugar molecule has more chemical energy than the carbon dioxide and water molecules that were used in its creation.

In our diet, we measure this energy in *kilocalories*. A calorie is the energy necessary to raise one gram of water one degree Celsius: it is about 4.19 joules. This is a very small unit: an apple has about 100,000 calories( 100 kilocalories), so people working with food started measuring everything in kilocalories.

Here is where things get confusing: People who work with food got tired of saying "kilocalories", so they just started using "Calorie" to mean 1,000 calories. This has created terrible confusion over the years. So if the C is capitalized, "Calorie" probably means kilocalorie.

### 1.1.4 Kinetic Energy

A mass in motion has energy. For example, if you are in a moving car and you slam on the breaks, the energy from the motion of the car will be converted into heat in the breaks and under the tires.

How much energy does the car have?

#### Formula for Kinetic Energy

$$E = \frac{1}{2}mv^2$$

where E is the energy in joules, m is the mass in kilograms, and v is the speed in meters per second.

### 1.1.5 Gravitational Potential Energy

Watch Khan Academy's **Potential energy** at <https://youtu.be/oGzwVYPxKjg>

When you lift something heavy onto a shelf, you are giving it *potential energy*. The amount of energy that you transferred to it is proportional to its weight and the height that you lifted it.

On the surface of the earth, gravity will accelerate a heavy object downward at a rate of  $9.8\text{m/s}^2$ .

#### Formula for Gravitational Potential Energy

On earth, then, gravitational potential energy is given by

$$E = (9.8)mh$$

where E is the energy in joules, m is the mass of the object you lifted, and h is the height that you lifted it.

There are other kinds of potential energy. For example, when you draw a bow, you have given that bow potential energy. When you release it, the potential energy is transferred to the arrow, which expresses it as kinetic energy.

## 1.2 Conservation of Energy

The first law of thermodynamics says “Energy is neither created nor destroyed.”

Energy can change forms: Your cells consume chemical energy to give gravitational potential energy to a car you push up a hill. However, the total amount of energy in a closed system stays constant.

### Exercise 1 The Energy of Falling

A 5 kg cannonball falls off the top of a 3 meter ladder. Just before it hits the floor, all of its gravitational potential energy has been converted into kinetic energy. How fast is the cannonball going when it hits the floor?

*Working Space*

*Answer on Page ??*

## 1.3 Efficiency

Watch Khan Academy's **Laws of thermodynamics** at <https://www.khanacademy.org/science/ap-biology/cellular-energetics/cellular-energy/a/the-laws-of-thermodynamics>

Although energy is always conserved as it moves through different forms, scientists aren't always that good at controlling it.

For example, a car engine consumes the chemical energy in gasoline. Only about 20% of the energy consumed is used to turn the wheels. Most of the energy is actually lost as heat. If you run a car for a while, the engine gets very hot and the exhaust going out the tailpipe turns hot.

A human is about 25% efficient. Most of the loss is in the heat produced during the chemical reactions that turns food into motion.

In general, if you are trying to increase efficiency in any system, the solution is usually easy to identify because heat is produced. Reduce heat, Increase efficiency.

Light bulbs are an interesting case. To get the light of a 60 watt incandescent bulb, you can

use an 8 watt LED or a 16 watt fluorescent light. Thus, we say that the LED light is much more efficient: If you run both, the incandescent bulb will consume 1.44 kilowatt-hours. The LED will consume only 0.192 kilowatt-hours.

Besides light, the incandescent bulb is producing a lot of heat. If it is inside your house, what happens to the heat? It warms your house.

In the winter, when you want light and heat, the incandescent bulb is 100% efficient!

In the summer, if you are running the air conditioner, the incandescent bulb is worse than just “inefficient at making light” – it is actually counteracting the air conditioner!



## APPENDIX A

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# Answers to Exercises

### Answer to Exercise ?? (on page ??)

At the top of the ladder, the cannonball has  $(9.8)(5)(3) = 147$  joules of potential energy.

At the bottom, the kinetic energy  $\frac{1}{2}(5)v^2$  must be equal to 147 joules. So  $v^2 = \frac{294}{5}$ . Thus it is going about 7.7 meters per second.

(Yes, a tiny amount of energy is lost to air resistance. For a dense object moving at these relatively slow speeds, this energy is negligible.)







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