

test

Grade:

STUDENT ID



BEYER HIGH SCHOOL

2015-2016

For:

$$\frac{-1.65 \times 10^{-19} \text{ J} (1.6 \times 10^{-19} \text{ C}) (1.6 \times 10^{-19} \text{ C})}{(1.6 \times 10^{-19} \text{ C})^2} \left(1 - \frac{1}{2}\right)$$

$$E = -1.55 \times 10^{-19} \text{ J}$$

For the higher energy level due to the greater number of electrons.

