

Half equation practice

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| 1. | $\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}_2$ | 11. | $\text{Ba}^{2+} + \text{e}^- \rightarrow \text{Ba}_2$ |
| 2. | $\text{K}^+ \rightarrow 2\text{K} + \text{e}^-$ | 12. | $\text{Ba}^{2+} + \text{e}^- \rightarrow \text{Ba}_2$ |
| 3. | $\text{I}^- \rightarrow \text{I} + \text{e}^-$ | 13. | $\text{H}^+ \rightarrow \text{H} + \text{e}^-$ |
| 4. | $\text{Na}^+ \rightarrow \text{Na} + \text{e}^-$ | 14. | $\text{Cl}^- \rightarrow \text{Cl} + \text{e}^-$ |
| 5. | $\text{O}^{2-} \rightarrow \text{O}_2 + \text{e}^-$ | 15. | $\text{Ag}^+ \rightarrow 4\text{Ag} + \text{e}^-$ |
| 6. | $\text{Br}^- \rightarrow \text{Br} + \text{e}^-$ | 16. | $\text{Ba}^{2+} + \text{e}^- \rightarrow \text{Ba}_2$ |
| 7. | $\text{Zn}^{2+} + \text{e}^- \rightarrow 3\text{Zn}_2$ | 17. | $\text{Na}^+ \rightarrow 2\text{Na} + \text{e}^-$ |
| 8. | $\text{Cl}^- \rightarrow \text{Cl} + \text{e}^-$ | 18. | $\text{Li}^+ \rightarrow 3\text{Li} + \text{e}^-$ |
| 9. | $\text{I}^- \rightarrow \text{I} + \text{e}^-$ | 19. | $\text{Fe}^{2+} + \text{e}^- \rightarrow \text{Fe}_2$ |
| 10. | $\text{Ba}^{2+} + \text{e}^- \rightarrow \text{Ba}_2$ | 20. | $\text{Fe}^{2+} + \text{e}^- \rightarrow \text{Fe}_2$ |