

EXPERIMENT

Conductivity of electrolyte solutions

Goal

The terms of **conductivity** and **electrolytes** will be introduced and related to the concept of dissociation. The use of the conductivity meter, very similar to the pH meter, will help to differentiate between strong and weak electrolytes. In the last part of the experiment, the presence of chemical reactions will be identified by measuring changes in conductivity.

Materials

- | | |
|--|--|
| <input type="checkbox"/> jumbo test tubes | <input type="checkbox"/> 0.1M solutions of KNO ₃ , KCl, KOH, HCl, and Ca(NO ₃) ₂ |
| <input type="checkbox"/> conductivity meter | <input type="checkbox"/> an unknown solution |
| <input type="checkbox"/> 6M solution of NH ₃ | <input type="checkbox"/> 5mL, 10mL, 25mL pipets (preferred Mohr's) |
| <input type="checkbox"/> 6M solution of CH ₃ COOH | <input type="checkbox"/> 250mL and 50mL beakers |
| <input type="checkbox"/> Ethyl alcohol | <input type="checkbox"/> rinsing water bottle |

Background

Playing with water near electrical appliances can be risky. Surprisingly, however, water does not conduct electricity. For electricity to propagate through media there must be free charges that move along the electric field. Pure water does not contain a significant number of charges and does not conduct electricity. However, when there is something dissolved in the water, therefore making an aqueous solution, it can be conductive.

Conductivity and aqueous solutions

Whether the aqueous solution is conductive or not depends on the nature of the solute. If the solute dissociates into ions when dissolved in water, the ions will conduct electricity. These types of chemicals are called electrolytes. If the solute dissolves as a molecule without dissociation, the solution will not conduct electricity and the solute will be called nonelectrolyte.

Conductivity units

The unit for conductivity is Siemens per centimeter, S-cm⁻¹. Pay attention to the units during this experiment because the tool will often display prefixes like milliSiemens or microSiemens per centimeter. Temperature affects conductivity and must be recorded during the measurement of this property.

Types of electrolytes and factors affecting their conductivity.

Salt and sugar are soluble in water. When sugar dissolves in water its molecules remain intact because water can not break apart the strong covalent bonds between their atoms. As a neutral molecule, sugar in solution will not conduct electricity and it is an example of a **nonelectrolyte**. Salt, sodium chloride or NaCl, is an ionic compound that gives place to two ions, Na⁺ and Cl⁻ when dissolved in water. These ions are responsible for the solution's conductivity. The number or more specifically the concentration, and the mobility of these ions will determine how conductive the solution is. Based on these principles we can classify substances as:

- × **Non-electrolytes.** When the aqueous solution of that substance does not conduct electricity. In general terms, these electrolytes present low conductivities, lower than $10\mu\text{S}\cdot\text{cm}^{-1}$.
- × **Weak electrolyte.** When the aqueous solution of that substance conducts electricity poorly. In general terms, these electrolytes present medium conductivities, between $10\text{-}1000\mu\text{S}\cdot\text{cm}^{-1}$.
- × **Strong electrolyte.** When the conductivity of the aqueous solution of that substance is high. In general terms, these electrolytes present large conductivities, larger than $1000\mu\text{S}\cdot\text{cm}^{-1}$.

The **mobility** of the ions is related to their size. Small ions move fast which enhances conductivity, while large ions encounter more resistance to move and show lower conductivities. The table below displays a series of mobility values for positive and negative ions at infinite dilution and 298K and 1 atm.

Ion	H_3O^+	Li^+	Na^+	Mg^{2+}	OH^-	Cl^-	Br^-	NO_3^-
mobility, u^∞ ($10^5 \text{ cm}^2 \cdot \text{V}^{-1} \cdot \text{s}^{-1}$)	363	41	52	55	206	80	81	74

The effect of ions' mobility is easily appreciated in isomolar solutions, which are solutions with the same molarity. The more ions in the solution, the higher the conductivity will be. For the same substance, the number of ions is determined by the **concentration**; more concentration means more ions and more conductivity. It is important to recall that only ions contribute to conductivity. While some molecules dissociate completely, others dissociate only partially. The latter are weak electrolytes. Interestingly, when two solutions containing ions are mixed, the conductivity of the resulting solution will be equal to the sum of the two separate solutions with the same concentration as in the mixture. This is called the **additivity rule** and will hold as long as the ions do not react with each other. When two electrolytes undergo a chemical reaction, new products will form and the additivity rule will not apply. A large deviation from the sum of separate conductivities is an indication of a chemical reaction.

Example

The conductivity of a 0.050 M solutions of HNO_3 is measure to be 19.6 mS/cm. That of a 0.050 M solution of KCl is 9.6 mS/cm. A new solution is prepared mixing 10 mL of HNO_3 0.10 M and 10 mL of KCl 0.10 M. The conductivity of the resulting mixture is measure to be 29.6 mS/cm. Calculate the concentration of each spices in the mixture and determine if their compounds have reacted or not.

Answer: To calculate the new concentration use the formula for dilutions:

$$V_1 \times M_1 = V_2 \times M_2$$

where V_1 and M_1 are the initial volume and concentration and V_2 and M_2 are the values after the dilution. We look for M_2 , therefore

$$M_2 = V_1 \times \frac{M_1}{V_2}$$

for HNO_3 ,

$$M_2 = \frac{10 \text{ mL} \times 0.10 \text{ M}}{(10 \text{ mL} + 10 \text{ mL})} = 0.05 \text{ M}$$

The same values apply for KCl.

We can see that the addition of the separate conductivities (κ) is very close to the total conductivity of the mixture with the same concentrations for each electrolyte.

$$\begin{aligned}\kappa(\text{HNO}_3 \text{ 0.05 M}) + \kappa(\text{KCl 0.05 M}) &= 19.6 \text{ mS/cm} + 9.6 \text{ mS/cm} = 29.2 \text{ mS/cm} \\ \kappa(\text{HNO}_3 \text{ 0.05M} + \text{KCl 0.05M}) &= 29.6 \text{ mS/cm}\end{aligned}$$

The additivity rules applies, meaning that there is no reaction between HNO_3 and KCl.

Procedure



Part A. Getting started: standardize the conductimeter

- ☐ *Step 1:* – Install the conductivity meter. The probe of the instrument is immersed in a clean solution that must be kept aside until the end of the experiment. Mind the probe needs to be hydrated at all times.
- ☐ *Step 2:* – To calibrate the conductivity meter, place 40mL of standard conductivity solution in a 50mL beaker while placing the probe in it.
- ☐ *Step 3:* – Now press Setup (top right menu). Normal setups are: cell constant of 1.0, standard recognition with manual selected, reference temperature 25 degrees, temperature coefficient of 2.1%, and alarm limits off.
- ☐ *Step 4:* – Now press Standardize (top right menu) and input 0.999 mS/cm in the keypad and press confirm (this should be the conductivity value of the standard). Finally, press Confirm (top right menu) to save the standard value. Now you are ready to measure.

Part A. Getting started: measuring conductivity

- ☐ *Step 5:* – It is very important to follow some cleaning steps before you measure the conductivity of any solution. You will need a 100 mL beaker labeled as waste, a wash bottle, and a 250 mL beaker with distilled water.
 1. Rinse the probe using the wash bottle over the waste beaker.
 2. Dip the probe 3-4 times into the 250 mL beaker with clean distilled water.
 3. Dip the probe into the solution to be measured. Use a 50 mL beaker (or a jumbo test tube) with at least 25 mL of solution. Use the same volume for all conductivity measurements.
 4. Repeat cleaning steps 1 and 2.

Good Lab Practice

-  To ensure the probe is well cleaned and the water in the 250 mL beaker remains clean, the conductivity meter should display a value close to that of the distilled water during the cleaning.
-  If the conductivity value of the cleaning water differs from the one of distilled water, replace the water in the 250 mL beaker with fresh distilled water, and thoroughly clean the probe.

Part B. Pure substances.

- ☐ *Step 6:* – Following the procedure outlined above, measure the conductivity of distilled water. Record the value in the Results section.
- ☐ *Step 7:* – Following the procedure outlined above, measure the conductivity of tap water. Record the value in the Results section.
- ☐ *Step 8:* – Following the procedure outlined above, measure the conductivity of Ethyl alcohol (C₂H₅OH). Record the value in the Results section.
- ☐ *Step 9:* – Classify the substance as strong, weak, or nonelectrolyte.

Part C. Series of electrolytes.

- ☐ *Step 10:* – In a clean 50 mL beaker, add 20 mL of 0.10 M HCl and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section.
- ☐ *Step 11:* – In a clean 50 mL beaker, add 20 mL of 0.10 M KOH and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section.
- ☐ *Step 12:* – In a clean 50 mL beaker, add 20 mL of 0.10 M KCl and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section.
- ☐ *Step 13:* – In a clean 50 mL beaker, add 20 mL of 0.10 M KNO₃ and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section.
- ☐ *Step 14:* – In a clean 50 mL beaker, add 20 mL of 0.10 M Ca(NO₃)₂ and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section.
- ☐ *Step 15:* – Find an unknown solution. In a clean 50 mL beaker, add 20 mL of the unknown solution and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity.
- ☐ *Step 16:* – Classify all substances as strong, weak, or nonelectrolytes and identify the unknown compound based on the conductivity values. Know that the unknown can also be water.

Part D. Series of Dilutions.

- ☐ *Step 17:* – Measure the conductivity of a 0.1 M solution of HCl. Write down the result in the Results section.
- ☐ *Step 18:* – Now you are going to prepare two dilutions. Starting from the concentrated 0.1 M solution of HCl, you are going to prepare 40mL volumes of the following dilutions: 0.050 M and 0.020 M. The first is a $\frac{1}{2}$ dilution, that is 20ml of acid needs to be mixed with 20ml of water to get 40mL of 0.050M solution, whereas the second is a $\frac{1}{5}$ dilution, that is 8ml of acid needs to be mixed with 32ml of water. Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers.
- ☐ *Step 19:* – Measure the conductivity of the two diluted HCl solutions (0.050 M, and 0.020 M) and record the values.

Part E. Mixtures of HCl and KNO₃.

- ☐ *Step 20:* – Write down the conductivity of the 0.050 M solution of HCl from the previous section in the Results section.
- ☐ *Step 21:* – Write down the conductivity of the solution of KNO₃ from the previous section in the Results section.
- ☐ *Step 22:* – In a clean 50 mL beaker, add 20 mL of 0.10 M HCl and 20 mL of 0.10 M KNO₃. Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section, minding that both solutions have now the same concentration and therefore only one number needs to be listed.

Part F. Mixtures of NH_3 and $\text{HC}_2\text{H}_3\text{O}_2$.

- ☐ *Step 23:* – Prepare the following dilutions: 100mL of acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) 0.12 M from 6M (add 2mL of concentrated solution and fill up with water until 100mL) and 40mL of ammonia (NH_3) 0.12 M from 6M (add 2mL of concentrated solution and fill up with water until 100mL). Carefully read the concentration of the initial concentrated solutions indicated in the labels. Work in the hood when handling highly concentrated solutions.
- ☐ *Step 24:* – In a clean 50 mL beaker, add 20 mL of the 0.12 M NH_3 dilution and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section. Work in the hood when handling highly concentrated solutions.
- ☐ *Step 25:* – In a clean 50 mL beaker, add 20 mL of 0.12 M $\text{HC}_2\text{H}_3\text{O}_2$ and 20 mL of distilled water (this is a $\frac{1}{2}$ dilution). Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section.
- ☐ *Step 26:* – In a clean 50 mL beaker, add 20 mL of 0.12 M NH_3 and 20 mL of 0.12 M $\text{HC}_2\text{H}_3\text{O}_2$. Use a Mohrs pipet to transfer the liquids, making sure you rinse with distilled water between transfers. Mix well and measure the conductivity. Recalculate the concentration after the dilution and write it down in the Results section, minding that both solutions have now the same concentration and therefore only one number needs to be listed.

⚠ CAUTION!

- ⚠ Handle concentrated acid or bases with care to avoid chemical burns.
- ⚠ Are you wearing your goggles? Do you know where the eye-washer is?

STUDENT INFO

Name:

Date:

Pre-lab Questions

Conductivity of electrolyte solutions

1. Starting with a 6 M solution of NaCl, calculate the volume of solution and the volume of water necessary to prepare 10 mL of a 3 M solution of NaCl.
2. How can you dilute a 3 M solution to form a 1.5 M solution? (e.g. I would pick up X mL of the solution and YmL of distilled water.)
3. How can you prepare 40mL of acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) 0.12 M from 6M.
4. How can you prepare 40mL of ammonia (NH_3) 0.12 M from 6M.
5. Given the data below, classify the following electrolytes as strong, weak, or nonelectrolytes:

	Concentration (M)	Conductivity	Units	Type of electrolyte
CuSO_4	0.05M	4000	$\mu\text{S} \cdot \text{cm}^{-1}$	_____
FeCl_3	0.05M	10000	$\mu\text{S} \cdot \text{cm}^{-1}$	_____
$\text{C}_6\text{H}_{12}\text{O}_6$	0.05M	120	$\mu\text{S} \cdot \text{cm}^{-1}$	_____
CH_3COOH	0.05M	320	$\mu\text{S} \cdot \text{cm}^{-1}$	_____

STUDENT INFO

Name:

Date:

**Results
EXPERIMENT****Conductivity of electrolyte solutions****Part B: Pure substances**

	Conductivity	Units	Electrolyte type
Distilled water	_____	_____	_____
Tap water	_____	_____	_____
C ₂ H ₅ OH	_____	_____	_____

Part C: Electrolytes

	Concentration (M)	Conductivity	Units	Electrolyte type
KOH	_____	_____	_____	_____
KCl	_____	_____	_____	_____
KNO ₃	_____	_____	_____	_____
Ca(NO ₃) ₂	_____	_____	_____	_____
Unknown		_____	_____	
Unknown #	_____	Unknown identity	_____	

Part D: Concentration effect

	Concentration (M)	Conductivity	Units
HCl	_____	_____	_____
HCl	_____	_____	_____
HCl	_____	_____	_____

Part E: HCl+KNO₃ mixture

	Concentration (M)	Conductivity	Units
HCl	_____	_____	_____
KNO ₃	_____	_____	_____
HCl+KNO ₃	_____	_____	_____

Part F: HC₂H₃O₂+NH₃ mixture

	Concentration (M)	Conductivity	Units
HC ₂ H ₃ O ₂	_____	_____	_____
NH ₃	_____	_____	_____
HC ₂ H ₃ O ₂ +NH ₃	_____	_____	_____

STUDENT INFO

Name:

Date:

Post-lab Questions

Conductivity of electrolyte solutions

1. Based on your results on Part B, which conductivity was higher, the one for distilled water or for tap water? Why?
2. Based on your results on Part C, specifically the conductivity values of KCl and KOH, how does ion mobility affect conductivity (see table in the Background section)? Use your data to illustrate your answer.
3. Based on your results on Part D, how does concentration affect conductivity? Use your data to illustrate your answer.
4. Based on your results in Part E, determine if their compounds have reacted or not. Does the additivity rule hold for this case?
5. Based on your results in Part F, determine if their compounds have reacted or not. Does the additivity rule hold for this case?

EXPERIMENT

Softening of hard water

Goal

The goal of this experiment is to understand the concepts of hard and soft water and their implications in daily life. The student will also experiment with the *dissociation of ionic compounds* in water, from which the **hydrogen-donor** nature of the acids and the notion of *pH* are emphasized. Other principles, such as equilibrium and ion exchange, are also presented.

Materials

- | | |
|--|---|
| <input type="checkbox"/> 250mL and 50mL beakers | <input type="checkbox"/> $\text{NH}_3/\text{NH}_4\text{Cl}$ buffer |
| <input type="checkbox"/> cation-exchange resin | <input type="checkbox"/> 0.01M $\text{Na}_2\text{H}_2(\text{EDTA})$ |
| <input type="checkbox"/> 6M HCl | <input type="checkbox"/> indicator (Eriochrome black mixed with NaCl) |
| <input type="checkbox"/> distilled water (not tap water) | <input type="checkbox"/> dropper solutions (0.10 M $\text{Ca}(\text{NO}_3)_2$, 0.10 M Na_2CO_3 , and 0.10 M $\text{Mg}(\text{NO}_3)_2$) |
| <input type="checkbox"/> Litmus paper | <input type="checkbox"/> 0.10 M NaHCO_3 solution |
| <input type="checkbox"/> 50mL graduated cylinder | <input type="checkbox"/> solid CaCO_3 |
| <input type="checkbox"/> 125mL Erlenmeyer | |

Background

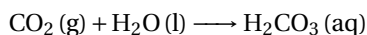
The term "**hard water**" refers to the concentration of certain minerals in the water. This topic concerns **aqueous solutions** in which water is the solvent and minerals are the solutes. These minerals are ionic compounds that dissociate in aqueous solutions producing ions. The hardness of water is an indication of the concentration of Ca^{2+} and Mg^{2+} ions.

Dissolving the insoluble

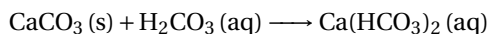
Ca^{2+} and Mg^{2+} ions originate in limestone (CaCO_3 , the material of the pyramids) and chalk (MgCO_3 , used to write on the chalkboard) deposits. However, chalk will not dissolve in water and neither the pyramids dissolve in the rain. How is dissolving these carbonates possible?

The answer is in the air

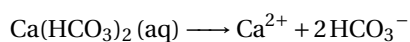
Water in contact with the air absorbs CO_2 in the atmosphere according to the equation:



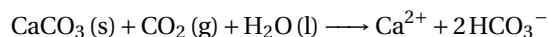
Where H_2CO_3 is carbonic acid, which reacts with the carbonate to form hydrogen bicarbonate. For the case of CaCO_3



Calcium hydrogen carbonate is also called calcium bicarbonate. Bicarbonates are soluble in water (baking soda is sodium bicarbonate, NaCO_3) and, as ionic compounds, they are dissociated into its ions;

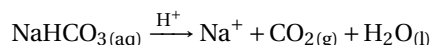


The overall reaction is the 3 previous equations added together:



and this is how carbonates dissolve in water.

Sodium hydrogen carbonate can decompose under different conditions. On one hand, in acidic conditions, it produces carbonic acid that quickly decomposes to give carbon dioxide:



On the other side, when heated it produces: $2\text{NaHCO}_3(\text{aq}) \xrightarrow{\Delta} \text{Na}_2\text{CO}_3 + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$

Implications in daily life

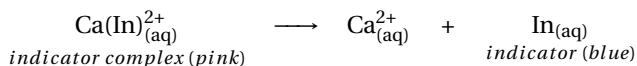
The presence of these ions in water has several implications for routine actions in our lives. The problem arises when the latter reaction takes place in the opposite direction, releasing CO_2 gas in the process and depositing solid carbonate. This direction is favored when water is heated and explains the white residue in bathrooms and kettles. This is especially a problem for boilers, that suffer from scaling when operating hard water.

The experiment

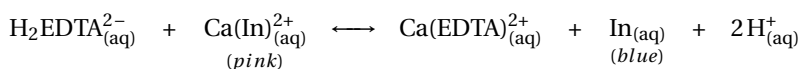
The experiment consists of four parts: preparing the ion exchange resin, measuring the hardness of a sample of water, reducing the hardness also called softening, of the sample, and testing the hardness of the softened sample. Additionally, a series of simple reactions will be done to test the reactivity of Ca^{2+} and Mg^{2+} ions.

Measuring the hardness of the sample

The concentration of Ca^{2+} ions, or any other metallic cation, is usually measured by titration of the aqueous sample using a sodium salt as the titrant, a solution of known concentration added to the sample. First, an indicator will be added to the sample which will color the solution pink. The indicator chosen is Eriochrome Black T, which will turn blue when the Ca^{2+} ion is removed from the cation-indicator complex:



To achieve the color change $\text{H}_2\text{EDTA}^{2-}$ will be used. $\text{H}_2\text{EDTA}^{2-}$ (dihydrogen ethylenediaminetetraacetate) is an anion with a stronger affinity for the Ca^{2+} cations than the indicator. As $\text{H}_2\text{EDTA}^{2-}$ is added to the solution it will replace the indicator forming a new complex, leaving the indicator alone, which in turn will become blue.



Even when the example described above refers to calcium ions, the same chemistry applies to magnesium ions. It is important to note down the exact number of drops required to achieve the endpoint of the titration.

Softening the hard water

After measuring the hardness of the water sample, the next step is to soften the water. *Softening* is the process where ions, particularly calcium and magnesium ions, are removed from the water. When a hard water sample is softened we can refer to it as *softened water*. A cation-exchange resin will be used to soften the water. This resin contains many sulfonic groups, ($-\text{SO}_3\text{H}$). Every two hydrogens in these acidic groups will be replaced by a metallic cation, Ca^{2+} and Mg^{2+} , in a mechanism called *cation exchange*. Finally, it is necessary to evaluate whether the water sample has been softened. As such the titration performed in Part A will be repeated, but this time with the softened sample. The most relevant parameter is the amount of EDTA added to achieve the endpoint of the titration, to compare this result to the previous titration. This experiment aims to be a qualitative analysis of the water treatment. For that reason, and to simplify the process, the titration will be done using droppers instead of the more precise burette typically used in titrations.

Procedure

Part A. Preparing the resin.

- ☐ *Step 1:* – Obtain around 5g of the dry, ion exchange resin in a 250 mL beaker. Remember that this resin can be recycled. Do not discard it at the end of the experiment.
- ☐ *Step 2:* – Add enough 6 M HCl to cover the resin. By adding the acid, the resin is cleaned of metallic cations. To use the ion exchange resin, it must be clean, also known as "in H⁺ form".

⚠ CAUTION!

- ⚠ Handle concentrated acid with care to avoid chemical burns.
- ⚠ Are you wearing your goggles? Do you know where the eye-washer is?

- ☐ *Step 3:* – Now you are going to wash the acid away. After 2 minutes, add 200 mL of distilled water to the resin-acid mixture. Do not move it, but rather allow the resin to settle in the bottom of the beaker.
- ☐ *Step 4:* – Carefully, pour the liquid off (decant) into another beaker. This liquid is an HCl solution and should be handled and discarded as such.
- ☐ *Step 5:* – To ensure that all the acid is gone, wash again the resin with 200 mL of distilled water. Allow the resin to settle in the bottom of the beaker and pour the liquid into another beaker.
- ☐ *Step 6:* – Before you discard the liquid you can test the pH. Dip the tip of a glass stirring rod in the liquid and poke the wet end of the rod into a piece of Blue Litmus paper. If the Litmus paper turns pink, there is still acid, repeat steps 5 and 6. If it does not turn pink, the resin is clean and ready to be used. Reserve the resin for Part B.

Part B. Measure the hardness of a sample of softened water

- ☐ *Step 1:* – Obtain a little bit more than 150 mL of hard water in a beaker. Hard water is already prepared in the lab.
- ☐ *Step 2:* – Using a graduated cylinder, add 20.0 mL of hard water to a 125 mL Erlenmeyer. Reserve this water for Part C. Use the rest of the water for this part of the experiment.
- ☐ *Step 3:* – Transfer the remaining hard water to the beaker with the resin and let it stand for 25 minutes. In the meantime, proceed with Parts C and D.
- ☐ *Step 4:* – Pour off about 40 mL of the water sample that has been in contact with the resin into a clean beaker. This sample would be referred to as the softened water.
- ☐ *Step 5:* – Using a graduated cylinder, add 20.0 mL of the softened water to a 125 mL Erlenmeyer.
- ☐ *Step 6:* – Add 5 mL of an NH₃/NH₄Cl buffer solution (not to be confused with Na₂H₂(EDTA)) to the Erlenmeyer and the tip of a spatula (less than a pea-sized amount) of the indicator (a solid mixture of eriochrome black T and sodium chloride). Adding too many indicators will ruin the experiment. The solution color will be rose pink after the addition of the indicator.
- ☐ *Step 7:* – Now, we are going to titrate the water sample. Obtain about 20 mL of 0.01 M Na₂H₂(EDTA) (the full name is Ethylenediaminetetraacetic acid) in a 50 mL beaker and a medicine dropper (or a plastic Pasteur pipette). Add the Na₂H₂(EDTA) drop by drop to the mixture while stirring. As the endpoint is approached the solution color will become lavender. Count the drops needed for the analyte to turn blue. When the addition of one drop turns the solution blue, all Ca₂⁺ and Mg₂⁺ ions have been removed from the sample
- ☐ *Step 8:* – Clean the resin as described in Part A, and return the resin to the instructor—there is a large beaker ready to collect wet resin.

Part C. Measure the hardness of a water sample

- ☐ *Step 1:* – While Part B takes place, and the water becomes softer, add 5 mL of a $\text{NH}_3/\text{NH}_4\text{Cl}$ buffer solution (not to be confused with $\text{Na}_2\text{H}_2(\text{EDTA})$) to the Erlenmeyer, and the tip of a spatula (less than a pea-sized amount) of the indicator. Remember that the indicator is just a solid mixture of eriochrome black T and sodium chloride. Adding too many indicators will ruin the experiment. The solution color will be rose pink after the addition of the indicator.
- ☐ *Step 2:* – Now, we are going to titrate the water sample. Obtain about 20 mL of 0.01 M $\text{Na}_2\text{H}_2(\text{EDTA})$ in a 50 mL beaker and a medicine dropper (or a plastic Pasteur pipette). Add the $\text{Na}_2\text{H}_2(\text{EDTA})$ drop by drop to the mixture while stirring. As the endpoint is approached the solution color will become lavender. Count the drops needed for the analyte to turn blue. When the addition of one drop turns the solution blue, all Ca^{2+} and Mg^{2+} ions have been removed from the sample

Part D. Observing reactions with Ca^{2+} and Mg^{2+} ions

- ☐ *Step 1:* – Obtain 4 test tubes and locate the following solutions: 0.10 M $\text{Ca}(\text{NO}_3)_2$, 0.10 M Na_2CO_3 , 0.10 M NaHCO_3 , and 0.10 M $\text{Mg}(\text{NO}_3)_2$.
- ☐ *Step 2:* – In test tube 1, add 10 drops of $\text{Ca}(\text{NO}_3)_2$. Add 10 drops of Na_2CO_3 and write down your observations.
- ☐ *Step 3:* – In test tube 2, add 10 drops of $\text{Ca}(\text{NO}_3)_2$. Add 2 drops of HCl and write down your observations.
- ☐ *Step 4:* – In test tube 3, add 10 drops of $\text{Ca}(\text{NO}_3)_2$. Add 10 drops of 0.1 M NaHCO_3 and write down your observations.
- ☐ *Step 5:* – Now, Heat test tube 3 in the flame of a Bunsen burner. Write down your observations.
- ☐ *Step 6:* – Now, cool down test tube 3. Add 2 drops of 6M HCl and write down your observations.
- ☐ *Step 7:* – In test tube 4, add a pea-sized amount of solid CaCO_3 . Add 2 drops of 6M HCl and write down your observations.
- ☐ *Step 8:* – Now, repeat the steps above using $\text{Mg}(\text{NO}_3)_2$ instead of $\text{Ca}(\text{NO}_3)_2$.
- ☐ *Step 9:* – At this point, complete the remaining steps of Part B.

STUDENT INFO

Name:

Date:

Pre-lab Questions

Softening of hard water

1. A student needs 10 drops of EDTA to reach the equivalency point when titrating a sample of hard water. After the students soften the water with the resin, it requires 12 drops of EDTA to reach the equivalency. Are these results correct? Explain.
2. Find the definition of the following concepts:
 - (a) Titration
 - (b) Titrant
 - (c) Litmus paper
 - (d) Ion exchange resin
3. Where is the term **hard water** derived from?
4. Compare the sizes used in the experiment. Circle the bigger size:
 - (a) Pea-size
 - (b) Tip of a spatula

STUDENT INFO

Name:

Date:

**Results
EXPERIMENT**

Softening of hard water

Parts B and C: Comparing titrations

Indicate how many drops you used to achieve the endpoint for:

Hard water sample (Part C): _____ drops of $\text{Na}_2\text{H}_2(\text{EDTA})$ Softened water sample (Part B): _____ drops of $\text{Na}_2\text{H}_2(\text{EDTA})$ **Part D: Observing reactions with Ca^{2+} and Mg^{2+} ions**

Write down your observations such as gas evolution, the appearance of precipitate, dissolving of precipitate, and no observation.

	$\text{Ca}(\text{NO}_3)_2$	$\text{Mg}(\text{NO}_3)_2$
Na_2CO_3		
HCl		
NaHCO_3		
$\text{NaHCO}_3 + \text{heat}$		
$\text{NaHCO}_3 + \text{heat} + \text{HCl}$		
Solid $\text{CaCO}_3 + \text{HCl}$		

STUDENT INFO

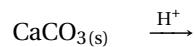
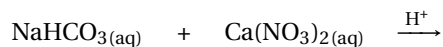
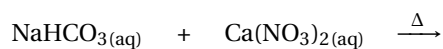
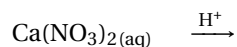
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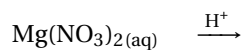
Post-lab Questions**Softening of hard water**

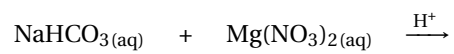
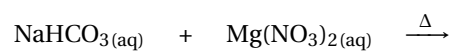
1. Based on your measurements. Did the ion exchange resin soften the hard water sample? Explain.

2. Write the balanced equations for the chemical reactions below. Mind that calcium carbonate is an insoluble compound.



3. Write the balanced equations for the chemical reactions below. Mind that magnesium carbonate is an insoluble compound.





4. Using the equivalency between drops and mL (1mL=15drops), calculate the number of mL of $\text{Na}_2\text{H}_2(\text{EDTA})$ needed for the titration of the soft and hard water samples.