1. ⬖ **STUDY CHECK**

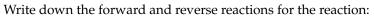Write down the forward and reverse reactions for the reaction:

$$CO(g) + O_2(g) \rightleftharpoons CO_2(g)$$

Show work:

Write just the final answer here:

2. ⬖ **STUDY CHECK**

The value of K_c for $F_2(g) \rightleftharpoons 2\,F(g)$ at 500K is 7×10^{-13}. Indicate whether the equilibrium mixture will contain mostly reactants, mostly products of both.
Show work:

Write just the final answer here:

3. ◈ **STUDY CHECK**

Write down the expression of K_c for the following reaction:

$Pb^{2+}_{(Aq)} + 2\,I^-_{(Aq)} \rightleftharpoons PbI_2(s)$. Show work:

Write just the final answer here:

4. ◈ **STUDY CHECK**

For the following reaction

$$SO_2Cl_{2(g)} \rightleftharpoons SO_{2(g)} + Cl_{2\,(g)}$$

The value of K_p at 300K is 0.05. Calculate the value of K_c at the same temperature. Show work:

Write just the final answer here: