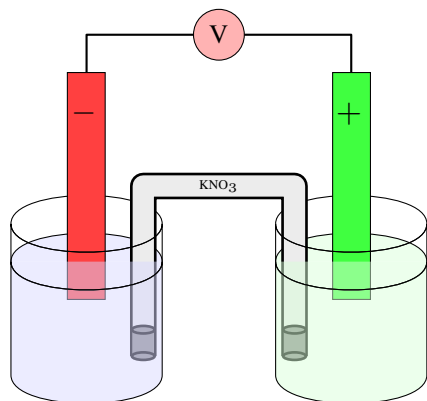


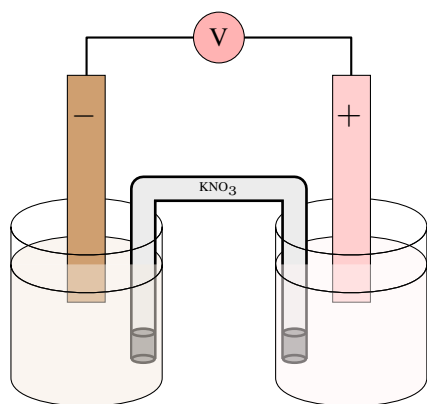
CHAPTER 0

INTRODUCTION TO GALVANIC CELLS

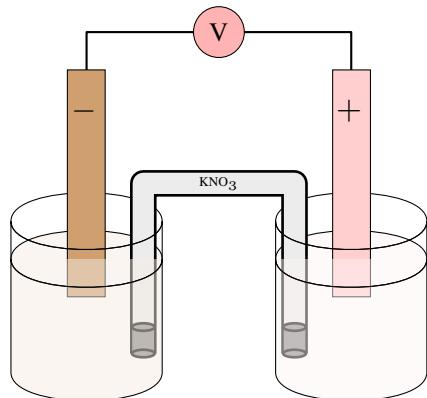
0.1 For the galvanic cell below, indicate: (a) the direction of flow of electrons (b) the direction of flow of cations (c) the direction of flow of anions



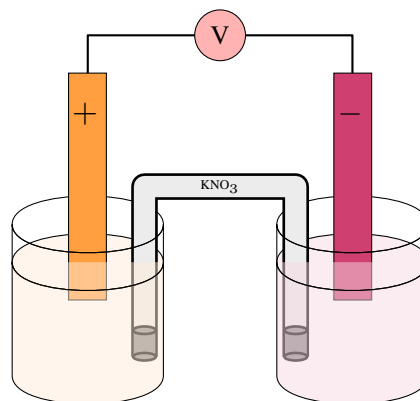
0.2 For the galvanic cell below, indicate: (a) the direction of flow of electrons (b) the direction of flow of cations (c) the direction of flow of anions



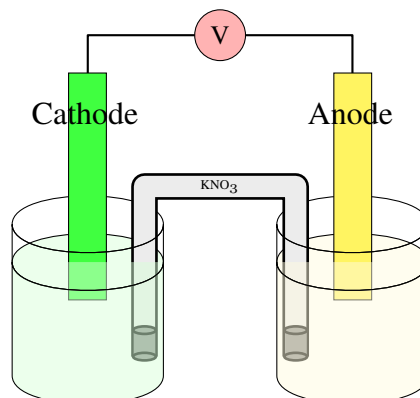
0.3 For the galvanic cell below, indicate: (a) label the anode (b) label the cathode



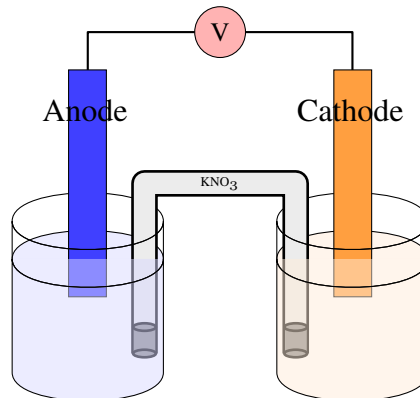
0.4 For the galvanic cell below, indicate: (a) label the anode (b) label the cathode



0.5 For the galvanic cell below, indicate: (a) label the sign (– or +) of each electrode (b) identify the flow of electrons

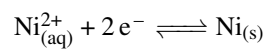


0.6 For the galvanic cell below, indicate: (a) label the sign (– or +) of each electrode (b) identify the flow of electrons

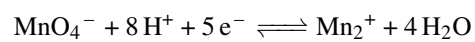


STANDARD REDUCTION POTENTIALS

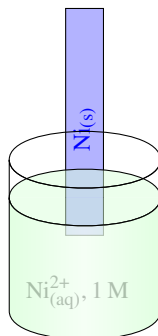
0.7 Sketch a semi-cell for the semi reaction below, indicate if the electrode is a solid-liquid electrode or gas-liquid electrode. Assume standard conditions:



0.8 Sketch a semi-cell for the semi reaction below, indicate if the electrode is a solid-liquid electrode or gas-liquid electrode. Assume standard conditions:



Answers **0.1** (a) left to right (b) left to left side of salt bridge (c) right to right side of salt bridge **0.2** (a) right to left (b) right to right side of salt bridge (c) left to left side of salt bridge **0.3** (a) left (b) right **0.4** (a) right (b) left **0.5** (a) left + , right – (b) from right to left **0.6** (a) right + , left – (b) from left to right **0.7** A solid-liquid electrode



0.8 A solid-liquid electrode

