

1

Organic Chemistry



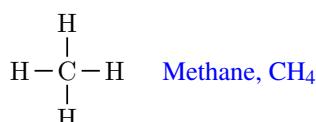
THE previous chapter have covered the chemical properties of what we call inorganic compounds. These are for example cooking salt (NaCl), water (H_2O) or ammonia (NH_3). This chapter covers organic compounds, which are chemicals based on carbon and hydrogen. We used organic compounds in the form of numerous materials such as fuels, perfumes, plastics or drugs. They all have common properties, from their smells to their powerful action. We will address first the naming of organic compounds as well as the properties of these important chemicals. Finally we will address the different groups of functional atoms that give unique properties to molecules such as caffeine or cocaine.

1.1 Alkanes

This first section will introduce organic chemistry, covering the most simple organic compounds: the alkanes. Alkanes simply contain carbon and hydrogen and are also called hydrocarbons. First you will be introduced to a few organic chemicals without paying to much attention to their names. Then, you will learn about a series of different organic formulas that can represent the same compound. Finally, you will learn the naming rule of alkanes, which basically extend to other-more complex-organic chemicals. ?

Identifying organic compounds How do we know when we have an organic or an inorganic chemical? Organic chemicals are based on carbon and hydrogen. They can contain other nonmetallic elements such as oxygen, nitrogen or sulfur. Here a few examples: CH_4 , C_2H_6 , C_6H_6 or CH_3COOH . Mind that carbonates (Na_2CO_3), carbon monoxide (CO) or carbon dioxide (CO_2) are not organic compounds.

Methane The simplest organic compound is methane: CH_4 . Methane is a fuel and the main constituent of natural gas. Going back a few chapters, the Lewis structure of methane is:



This structure is very representative in this chapter as it shows that each carbon atom has to be connected to four different atoms.

Alkanes Alkanes are simple organic compounds made of Carbon and Hydrogen with all carbons connected by means of simple bonds—these are single lines to represent the connections between atoms. The naming of alkanes results of adding a prefix to

GOALS

- 1 Identify organic chemicals
- 2 Name linear hydrocarbons
- 3 Name cyclic hydrocarbons
- 4 Name hydrocarbons with substituents
- 5 Identify functional groups

Discussion: Think about any drug you have taken recently. Paste its structure and indicate at least one functional group in the molecule.



Figure 1.1: Methane (CH_4) is used as a fuel for ovens, homes, water heaters.

# Carbons	prefix
1	Meth
2	Eth
3	Pro
4	But
5	Pent
6	Hex
7	Hepta
8	Octa
9	Nona
10	Deca

Figure 1.2: Prefixed for alkane naming.

Add Formula ?? to your
flashcard.

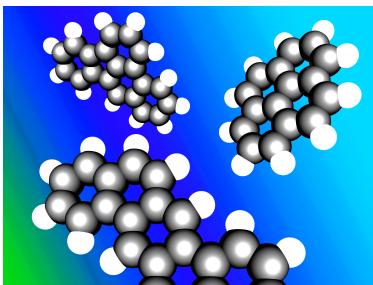
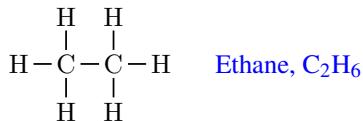
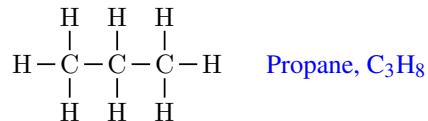


Figure 1.3: Hydrocarbons are made of carbon and hydrogen.

the suffix *ane*. The prefix depends on the number of carbons in the molecule. Table ?? shows a list of the different prefixes. For example, the alkane with one carbon is called methane. Another example of alkanes would be ethane that contains two carbons:



or propane:



General formula for alkanes At this point we saw two different alkanes: methane (CH_4), butane (C_2H_6) and propane (C_3H_8). The general formula for an alkane with n carbon atoms is:

$$\text{C}_n\text{H}_{2n+2} \quad (1.1)$$

As an example, the formula for methane ($n = 1$) is CH_4 and the formula for octane ($n = 8$) is C_8H_{18} .

Sample Problem 1

Write down the molecular formula for decane and pentane.

SOLUTION

Using Equation ?? we have that the molecular formula for decane ($n = 10$) would be: $\text{C}_{10}\text{H}_{22}$. Similarly, the molecular formula for pentane ($n = 5$) would be: C_5H_{12} .

◆ STUDY CHECK

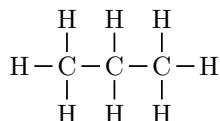
Name the alkane with formula C_7H_{16} and give the formula for nonane.

Answer: heptane, C_9H_{20} .

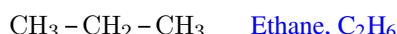


Figure 1.4: The octane rating of gas is a standard measure of the performance of engine fuels, originally determined by mixing a gasoline made entirely of heptane and 2,2,4-trimethylpentane (a highly branched octane).

 Alkanes contain CH_3 and CH_2 units Let us analyze the formula of propane. In this formula there are two different types of carbons. One is the end of the chain carbon, in the left and in the right. The other type of carbon is the central carbon. The extremes are bounded to three hydrogen, whereas the central is bounded to two hydrogens.



The extremes are indeed CH_3 units and the center is a CH_2 unit. So, another way to represent ethane would be:



At this point you have seen three different ways to represent organic molecules. Let us use propane as an example. One is the *molecular formula*, which in the case of propane is C_3H_8 . Another different way to represent propane is with its *expanded structural formula*, that is by representing all C and all H in the molecule. Another molecular representation is the *condensed structural formula*, that is by using CH_3 and CH_2 units. Here the three formulas:

C_3H_8	$\text{CH}_3-\text{CH}_2-\text{CH}_3$	$\begin{array}{c} \text{H} & \text{H} & \text{H} \\ & & \\ \text{H}-\text{C} & -\text{C} & -\text{C}-\text{H} \\ & & \\ \text{H} & \text{H} & \text{H} \end{array}$
molecular formula	condensed formula	expanded structural formula

Number of Carbons	Prefix	Condensed Structural formula
1	Meth	CH_4
2	Eth	CH_3-CH_3
3	Prop	$\text{CH}_3-\text{CH}_2-\text{CH}_3$
4	But	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_3$
5	Pent	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_3$
6	Hex	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_3$
7	Hepta	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_3$
8	Oct	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_3$
9	Non	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{CH}_3$

Sample Problem 2

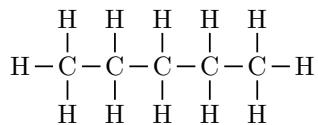
Write down the condensed and expanded formulas for pentane.

SOLUTION

Pentane has five carbons, hence its condensed formula will have two CH_3 units and three CH_2 units:

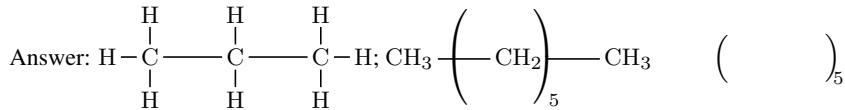


The expanded formula for pentane would be:



 **STUDY CHECK**

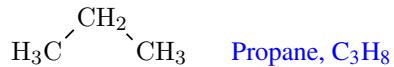
Write down the condensed and expanded formulas for heptane.



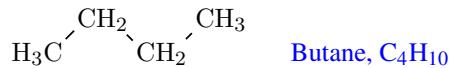
 *ZigZag* Let us analyze the structure of propane again. This molecule has three carbons in the form of a C-C chain such as: C-C-C. However, C-C chains with more than more three carbons are not linear. So instead of representing propane as:



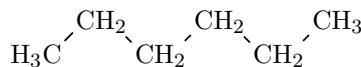
we could represent this molecule by means of a zigzag:



Another example would be butane:



 *Skeletal formula* A final way to represent molecules in this chapter is simply using lines. This is called the skeletal formula, as you only represent the C-C skeleton of the molecule. Let us use hexane an an example. We can represent this molecule as in this two ways:



condensed formula

skeletal formula

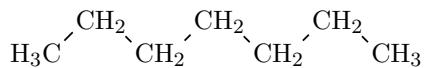
The ending of a skeletal formula represents a CH_3 and the points in between represents CH_2 .

Sample Problem 3

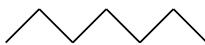
Write down the condensed and skeletal formulas for heptane.

SOLUTION

Heptane has seven carbons, hence its condensed formula will have two CH_3 units and five CH_2 units:



The skeletal formula for would be:



◆ STUDY CHECK

Draw the skeletal formula of decane.

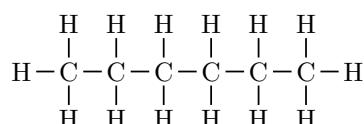
Answer:

1.2 Cycloalkanes

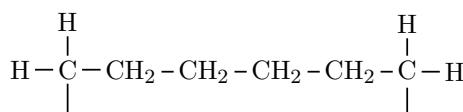
This section deals with cyclic alkanes. Alkanes are perfect examples of hydrocarbons, with C-C chains and all carbon atoms saturated with hydrogen. Cycloalkanes are in essence alkanes that have a cyclic structure. We will cover the molecular, condensed and skeletal formulas for these chemicals.



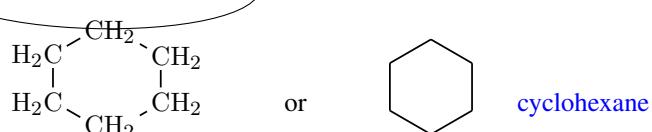
Cyclic alkanes Let us use the expanded structure of hexane as an example.



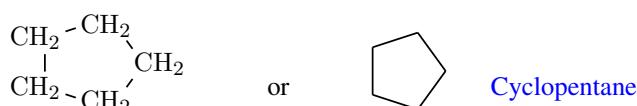
A cycloalkane results of eliminating the left and right hydrogen and connecting the molecule in the form of a cycle:



As the most stable structure for six lines is the hexagon. The results looks like



Naming cycloalkanes The naming of alkanes and cycloalkanes is very similar. You just need to add the cyclo prefix to the name. For example, the alkane with five carbons is called pentane. The corresponding cycloalkane is called cyclopentane:



Sample Problem 4

Write down the condensed structure and name the following cycloalkane:

SOLUTION

The cycloalkane has four carbons and its name is cyclobutane. Its condensed structure is

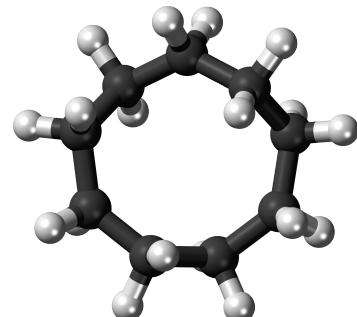


Figure 1.5: Cyclononane contains 9 carbons in a cycle.

Q: What is the molecular formula Cycalononane?

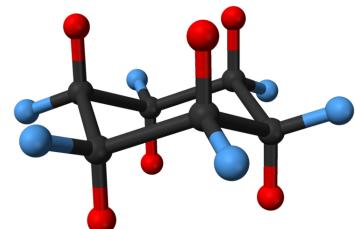


Figure 1.6: Cyclohexane has a chair shape.

Q: How many H are there in cyclohexane?

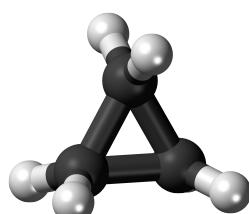
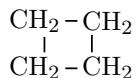


Figure 1.7: Cyclopropane has only three carbons and a lot tension in its cycle.

Q: Would cyclopropane be a stable or unstable molecule?



 **STUDY CHECK**

Write down the condensed structure and name the following cycloalkane:



Answer: cyclopropane, C_3H_6



General formula for cycloalkanes The general formula for an cycloalkane with n carbon atoms is:



(1.2)

Add Formula ?? to your
flashcard.

Sample Problem 5

Write down the molecular formula for cyclodecane and cyclopentane.

SOLUTION

Using Equation ?? we have that the molecular formula for cyclodecane ($n = 10$) would be: $\text{C}_{10}\text{H}_{20}$. Similarly, the molecular formula for cyclopentane ($n = 5$) would be: C_5H_{10} .

 **STUDY CHECK**

Name the alkane with formula C_7H_{14} and give the formula for cyclononane.

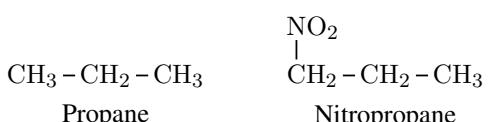
Answer: cycloheptane, C_7H_{16} .

Substituents	name
$\text{CH}_3 -$	Methyl
$\text{CH}_3\text{CH}_2 -$	Ethyl
$\text{CH}_3\text{CH}_2\text{CH}_2 -$	Propyl
$\begin{array}{c} \\ \text{H}_3\text{C} - \text{CH} - \text{CH}_3 \end{array}$	Isopropyl
$\begin{array}{c} \\ \text{H}_3\text{C} - \text{C} - \text{CH}_3 \\ \\ \text{CH}_3 \end{array}$	t-butyl
$\text{F} -$	Fluoro
$\text{Cl} -$	Chloro
$\text{Br} -$	Bromo
$\text{I} -$	Iodo
$\text{NH}_2 -$	Amino
$\text{NO}_2 -$	Nitro
$\text{CH}_2 = \text{CH} -$	Vinyl

Figure 1.8: Name of some substituents.

1.3 Alkanes with substituents

At this point you should be familiar with alkanes. These organic compounds are carbon-based chemicals made of lines of C-C hydrocarbon chains. Often times these hydrocarbons have other groups of atoms attached to the hydrocarbon chains and these are called substituents. This section cover the naming of alkanes with substituents. Here an example of an alkane and an alkane with a substituent:

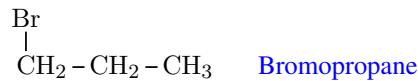


Substituents There are many different substituents which can be found connected to an alkane chain. Their names are indicated in Figure 1.8. The easiest substituents are halogens and atoms of chlorine ($\text{Cl} -$), bromine ($\text{Br} -$) or iodine ($\text{I} -$) can connect to an alkane. The name of these substituents is chloro, bromo and iodo. Other substituents can contain carbon, like a methyl ($\text{CH}_3 -$) or a ethyl ($\text{CH}_3\text{CH}_2 -$). There are even more complex substituents such as tert-butyl

 *Alkanes with a single substituent* Let us consider the following example. The condensed formula for propane is

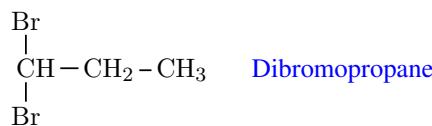


Now, this would be propane with a substituent:



As you can see a bromine atom substitutes one of the hydrogen atoms of the second first of the molecule (starting from the left).

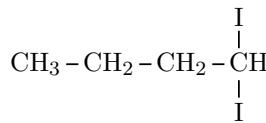
 *Alkanes with two or more equal substituents* In the same way as when you have a single bromine atom attached to propane, you can also have two Br —. In this case you need to use the prefix *di* to indicate there are two identical bromines. For example, the name of the following molecule would be



Similarly, you should use the prefix *tri* for three equal substituents.

Sample Problem 6

Name the following hydrocarbon:

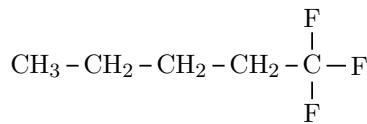


SOLUTION

The carbon chain has four carbons and hence the ending of the name would be: butane. Also there are two iodines (iodo substituents) attached to the carbon chain. As there are two of the same iodo atoms, we need to use the prefix *di*. The full name would be: Dioiodobutane.

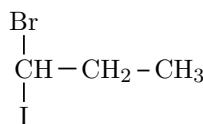
◆ STUDY CHECK

Name the following hydrocarbon:



Answer: Trifluoropentane.

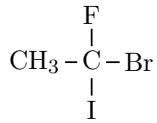
 *Alkanes with different substituents* Now imagine you have two different halogens as substituents: Br — and I — like in next example



As they have different names you cannot use the prefix *di*. Still, when you indicate the names of the substituents you need to order them according to the *abc*. So bromo goes first in the name and iodo after. You also need to separate the different substituents with a '-'. The final name of the hydrocarbon above would be: Bromo-Iodopropane.

Sample Problem 7

Name the following hydrocarbon:

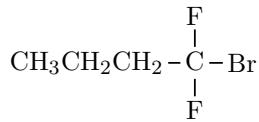


SOLUTION

The carbon chain has two carbons and hence the ending of the name would be: ethane. Also there are three different substituents: iodine (ido substituents), bromine (bromo substituents) and fluorine (fluoro substituents). We need to order them according to the *abc*, hence the order would be: bromo, then fluoro and finally iodo. The full name of the alkane would be: Bromo-Fluoro-Iodoethane.

 **STUDY CHECK**

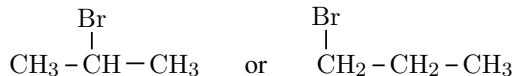
Name the following hydrocarbon:



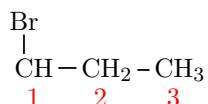
Answer: Bromodifluorobutane.



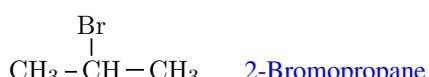
Numbering the chain Substituents are atoms or groups of atoms that can plug into a alkane chain. You could envision plugging these atoms at different points of the chain. For example:



In the right example Br is plugged to the left C atom, whereas in left example C is plugged to the middle carbon. Hence, it is important first, to learn how to number a hydrocarbon chain. Let us use propane as an example. This molecule has three atoms. In order to number the chain, you start by selecting the extreme that is the closest to the substituent, and use this carbon as number one. Next carbon would be carbon number two and so on until you arrive to carbon number three.

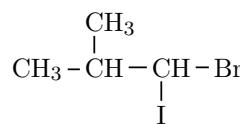


As the Br atom in the carbon number one, the name of the molecule would be: 1-bromopropane or simply bromopropane. Differently, when the substituent is in a carbon different than one, you need to indicate that location. For example:

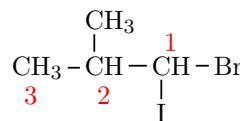


Sample Problem 8

Name the following hydrocarbon:

**SOLUTION**

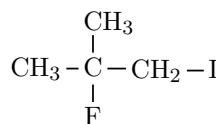
First we find the ending of the name: as the molecule has three carbons in the main chain, the ending of the name would be: propane. Then we need to number the chain so that the number one carbon is the closest to the substituents:



A methyl is connected to carbon two, and two halogens, a iodo and a bromo are connected to carbon number one. The substituents are: 2-methyl, 1-bromo, 1-iodo. If we order them: 1-bromo-1-iodo-2-methyl. And the final name would be: 1-bromo-1-iodo-2-methylpropane.

❖ STUDY CHECK

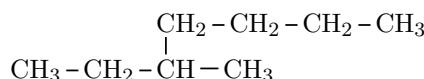
Name the following hydrocarbon:



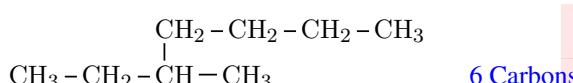
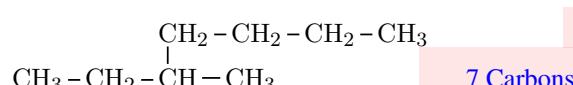
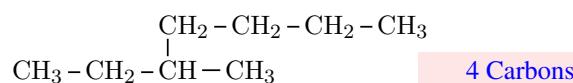
Answer: 1-iodo-2-fluoro-2-methylpropane.



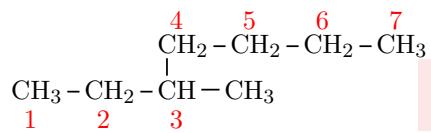
Finding the longest chain Alkanes with substituents are more complex than simple alkanes and often time they contain more than one hydrocarbon chain. Therefore, one can envision several ways to number the chain. The rule is to locate the longest chain. Let us use the following hydrocarbon. How many chains can you find, and which is the longest chain?



The answer should be three chains. Let me highlight the three different possibilities and the number of atoms in each chain:



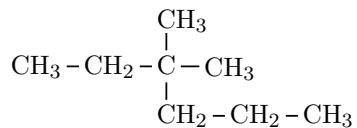
As the longest chain has seven carbons, the name of the molecule would be heptane. Still you need to add the substituents before that name. After you locate the longest chain you need to number the chain so that the substituents are located the closest to the carbon number one the possible:



So in carbon number three a methyl is located. Hence the final name of the molecule would be: 3-methylheptane.

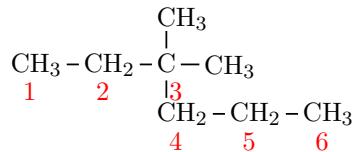
Sample Problem 9

Name the following hydrocarbon:



SOLUTION

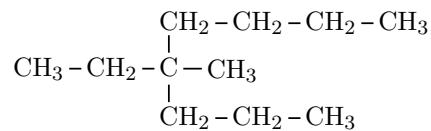
First we locate the longest chain. We have five possible chains, and the longest one has six carbons. Hence the name of the hydrocarbon would be hexane. Now we need to number the carbons so that we start numbering the closest to the substituents the possible.



We have two methyl connected to carbon number three. Hence the final name will be: 3-dimethylhexane.

STUDY CHECK

Name the following hydrocarbon:



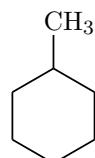
Answer: 4-ethyl-4-methyloctane.

1.4 Cylcoalkanes with substituents

Cycloalkanes are cyclic alkanes. This section covers the naming of cycloalkanes with substituents. The naming rules are the same as the rules for naming alkanes. This means first you will find the ending on the name by counting the number of carbons. Then you will locate each substituent and number the carbon chain so that these substituents are close to carbon number one. All substituents have a number depending on the carbon number. Finally, all substituents should be order according to the *abc*.

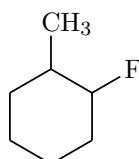


 *Cycloalkanes with one substituent* Let us take a look at the following cycloalkane:

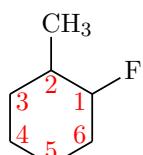


this is a cyclohexane connected to a methyl substituent. As there is only one substituents, there is no need to number the carbon chain. Hence the name would be: methylcyclohexane.

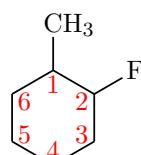
Cycloalkanes with two substituents Let us take a look at the following cycloalkane:



this is a cyclopentane connected to two different substituents: a methyl and a fluoro. In order to name this molecule we need to number the carbons first, and there are two different ways to number the cyclohexane ring:



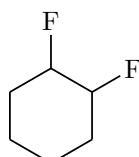
1-Fluoro-2-methylcyclopentane



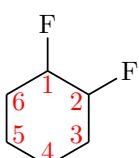
(wrong) Fluoro-1-methylcyclopentane

we will choose the name that gives the lowest numbers: 1-Fluoro-2-methylcyclopentane.

Cycloalkanes with repeated substituents Let us take a look at the following cycloalkane, which has two repeated fluorine substituents:



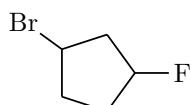
After numbering the chain:



The name would be: 1,2-difluorocyclohexane.

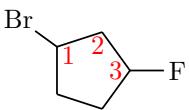
Sample Problem 10

Name the following hydrocarbon:



SOLUTION

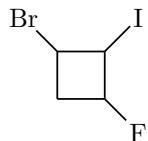
The molecule is a cyclopentane with two substituents: fluoro and bromo. I will start numbering in bromo and continue until bromo. This way I will have small numbers and follow the abc rule:



And the name of the molecule would be: 1-bromo-3-fluorocyclopentane.

◆ STUDY CHECK

Name the following hydrocarbon:



Answer: 1-bromo-3-fluoro-2-bromocyclopropane.

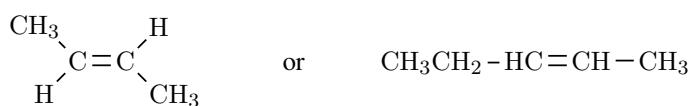
Functional groups	name
$\begin{array}{c} R_1 \backslash \\ \quad C=C \\ / \quad \backslash \\ R_2 \quad R_3 \\ \quad \\ R_4 \end{array}$	Alkene
$R - C \equiv C - R'$	Alkyne
$R - OH$	Alcohol
$R - SH$	Thiol
$R - O - R'$	Ether
$\begin{array}{c} O \\ \\ R - C - R' \end{array}$	Ketone
$\begin{array}{c} O \\ \\ R - C - H \end{array}$	Aldehyde
$\begin{array}{c} O \\ \\ R - C - OH \end{array}$	Carboxylic acid
$\begin{array}{c} O \\ \\ R - C - O - R' \end{array}$	Ester
$\begin{array}{c} R' \\ \\ R - N - R'' \end{array}$	Amine
$\begin{array}{c} O \quad R'' \\ \quad \\ R - C - N - R' \end{array}$	Amide

Figure 1.9: Name of functional groups.

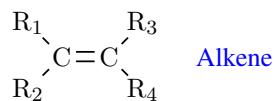
1.5 Molecular diversity

The previous sections covered hydrocarbons, either alkanes or cycloalkanes. Both alkanes and cycloalkanes are molecules based on carbon and hydrogen. This section will introduce the idea of functional group. You have certainly taken painkillers for a headache or over-the-counter drugs to get over a cold. Maybe you drink coffee and perhaps you like tea. All these medications as well as drinks contain active organic molecules. These active molecules differ from plain hydrocarbons, which are simply made of carbon and hydrogen. Active molecules contain functional groups such as alcohol, ethers, carboxylic acids, amines, amides or aromatic groups. These groups of atoms have a specific function and give activity to the molecule. The goal of this section is not for you to identify the action mechanisms of these complex active molecules, but simply to identify the different groups.

◆ **Alkene group: double bonds** Alkenes contain at least one double bond between carbons. An example would be:

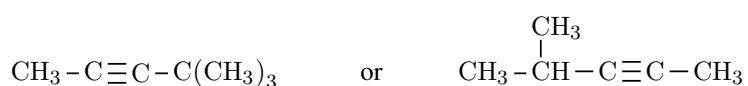


As both sides of an alkene can be connected to different hydrocarbon chains we normally represent this as:

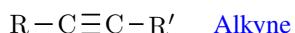


where R and R' represents any hydrocarbon chain.

◆ **Alkyne group: triple bonds** Alkenes contain at least one triple bond between carbons. An example would be:

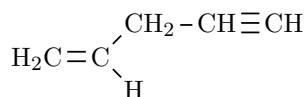


Again we could use R and R' to represent any hydrocarbon chain:

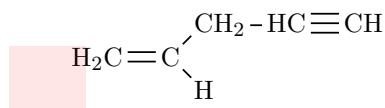


Sample Problem 11

Identify the alkene and alkyne groups in the molecule:

**SOLUTION**

A double bond is carbon atoms sharing two pairs of electrons, whereas a triple bond is a pair of atoms sharing three pairs of electrons. They are represented with a double and triple line, respectively. In the question:

**◆ STUDY CHECK**

Identify the alkene and alkyne groups in the molecule:

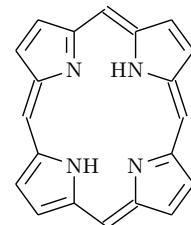
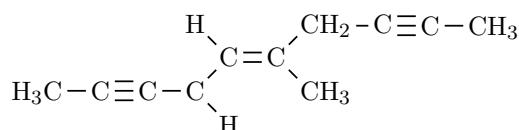
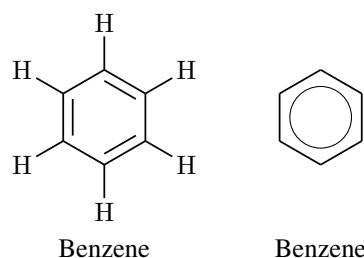


Figure 1.10: Porphyrin has several amine groups.

Q: How many amine groups has porphyrin?



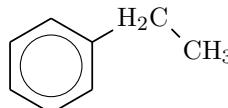
Aromatic group Aromatic groups are based on benzene, a cyclohexane with a series of alternating double bonds, which is often represented as a circle:



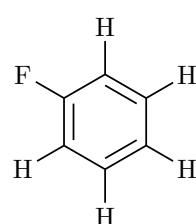
Examples of molecules containing aromatic groups are:



or



and



Alcohol, ether and thiols group Alcohols contain an $-OH$ group attached to a carbon.



Whereas ethers have oxygen atoms attached to two carbon atoms:



Examples of alcohols and ether are:

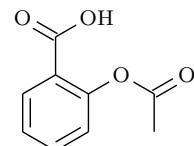


Figure 1.11: Aspirin contains an aromatic cycle, a carboxylic acid and a ester.

Q: What is the molecular formula of aspirin?

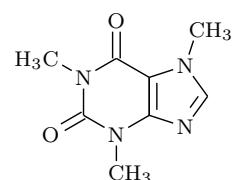


Figure 1.12: Caffeine contains amides and amines

Q: How many different amines and amides are there in caffeine?

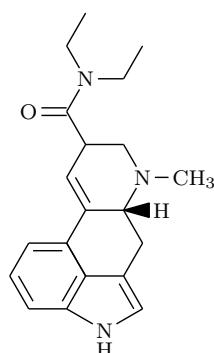


Figure 1.13: LSD has both an amide and an amine

Q: How many different amines has LSD?



alcohol

ether

Thiols contain a $- \text{SH}$ group attached to a carbon. They are equivalent to alcohols but based in sulfur:

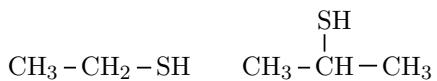


Figure 1.14: Esters give flavor to many fruits.
Q: What is the formula for the smallest ester?



Figure 1.15: Vinegar contains acetic acid, the smallest carboxylic acid.
Q: What is the formula for acetic acid?

Examples of thiols are:



Sample Problem 12

Classify the following molecules as alcohol or ether.



SOLUTION

The OH groups is an alcohol, and we find this group in the left molecule. Differently, the right molecule is an ether as it contains the $\text{R} - \text{O} - \text{R}'$ group.

❖ STUDY CHECK

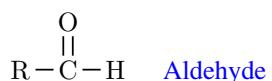
Classify the following molecules as alcohol or ether.



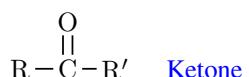
Answer: (left) ether; (right) alcohol.



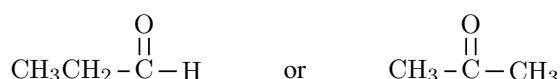
Aldehydes and ketones Ketones and aldehydes both contain a $\text{C}=\text{O}$ group. Still these are two different groups and ketones have a $\text{C}=\text{O}$ group bounded to two different carbon atoms



whereas aldehydes have the same $\text{C}=\text{O}$ group but this time bounded to a carbon and a hydrogen.



Examples of aldehydes and ketones are:

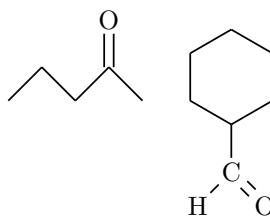


Aldehyde

Ether

Sample Problem 13

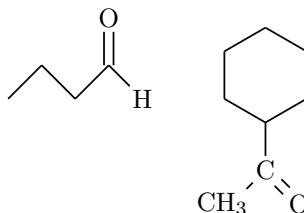
Classify the following molecules as an aldehyde or ketone

**SOLUTION**

The left molecule is a ketone as the carbonyl group ($\text{C}=\text{O}$) is connected to a CH_3 and a CH_2 . Differently, the right molecule is a ketone as the carbonyl group is connected to a cycloalkane but also to a hydrogen.

◆ STUDY CHECK

Classify the following molecules as an aldehyde or ketone



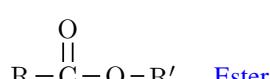
Answer: (left) aldehyde; (right) ketone.



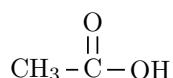
Carboxylic acids and esters Carboxylic acids contain a carbonyl group ($\text{C}=\text{O}$) connected to an hydrocarbon and also an alcohol group:



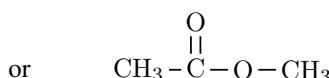
Esters have the same $\text{C}=\text{O}$ group but this time bounded to a carbon (R) and an ether group ($-\text{O}-\text{R}'$).



Examples of carboxylic acids and esters are:



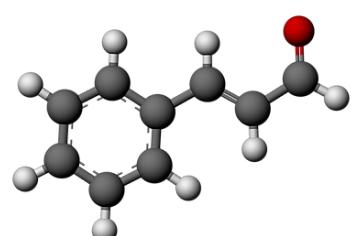
Carboxylic acid



Esters



Figure 1.16: Diethyl ether was formerly used as a general anesthetic, until non-flammable drugs were developed. The image shows a panel from monument in Boston commemorating the demonstration of ether's anesthetic use.



Sample Problem 14

Classify the following molecules as an carboxylic acid or ester:

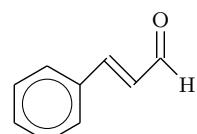


Figure 1.17: Cinnamaldehyde is the aldehyde that gives cinnamon its flavor and odor.



Figure 1.18: Benzyl acetate is an ester that flavors pears.

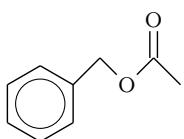
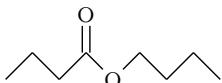


Figure 1.19: Butyl butyrate is an ester that flavors pineapple.



SOLUTION

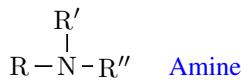
The molecule in the left is a carboxylic acid as the carbonyl group ($\text{C}=\text{O}$) is connected to an alcohol $-\text{OH}$. Differently, the molecule in the right is an ester as the carbonyl group is connected to a $-\text{O}-\text{CH}_3$ group.

◆ STUDY CHECK

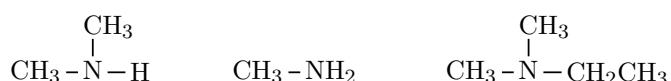
Classify the following molecules as an carboxylic acid or ester:

Answer: (left) ester; (right) acid.

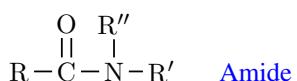
Amines and amides Amines and amides are groups containing nitrogen. Amines are derivative of ammonia (NH_3) with one or more of the hydrogen atoms being replaced by a hydrocarbon



All these molecules are amines:



Amides, on the other hand, contain a carbonyl group ($\text{C}=\text{O}$) connected to an amine group (—N—)

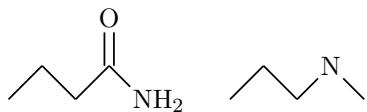


Examples of amides are:



Sample Problem 15

Classify the following molecules as an amide or amine:

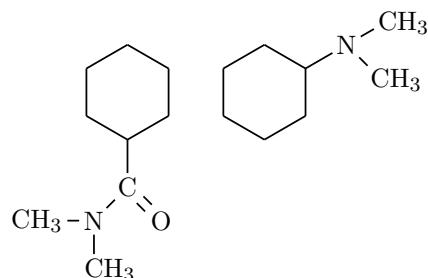


SOLUTION

The molecule in the left is a amide as the carbonyl group ($\text{C}=\text{O}$) is connected to a nitrogen atom. Differently, the molecule in the right is an amine as the nitrogen group is not connected to any carbonyl group.

◆ STUDY CHECK

Classify the following molecules as an amide or amine:



Answer: (left) amide; (right) amine.

CHAPTER 1

ALKANES

1. Classify the following chemicals in two different categories. Give a rationale for your classification. For the group with more chemicals, further classify those chemicals in two categories. (a) KCl (b) C₂H₂ (c) C₄H₁₀ (d) FeO_(s) (e) C₆H₁₂ (f) PH₃ (g) H₂O

Ans: (a) KCl (ionic, inorganic) (b) C₂H₂ (covalent, organic) (c) C₄H₁₀ (covalent, organic) (d) FeO_(s) (ionic, inorganic) (e) C₆H₁₂ (covalent, organic) (f) PH₃ (covalent, inorganic) (g) H₂O (covalent, inorganic)

2. Working in groups, select an everyday-life object (e.g. a spoon) and guess whether the materials that made this object are mostly organic or inorganic (e.g. mostly inorganic). Without revealing your answer, present the material you selected to another team member and ask him or her to give you his or her point of view regarding whether the materials that made this object are mostly organic or inorganic.

3. The molecular formula of ethane is

- | | |
|------------------------------------|-------------------------------------|
| (a) C ₂ H ₆ | (d) C ₆ H ₁₄ |
| (b) CH ₄ | |
| (c) C ₄ H ₁₀ | (e) C ₁₀ H ₂₂ |

Ans: (a)

4. The molecular formula of butane is

- | | |
|------------------------------------|-------------------------------------|
| (a) C ₂ H ₆ | (d) C ₆ H ₁₄ |
| (b) CH ₄ | |
| (c) C ₄ H ₁₀ | (e) C ₁₀ H ₂₂ |

Ans: (c)

5. The molecular formula of methane is

- | | |
|------------------------------------|-------------------------------------|
| (a) C ₂ H ₆ | (d) C ₆ H ₁₄ |
| (b) CH ₄ | |
| (c) C ₄ H ₁₀ | (e) C ₁₀ H ₂₂ |

Ans: (b)

6. The molecular formula of decane is

- | | |
|------------------------------------|-------------------------------------|
| (a) C ₂ H ₆ | (d) C ₆ H ₁₄ |
| (b) CH ₄ | |
| (c) C ₄ H ₁₀ | (e) C ₁₀ H ₂₂ |

Ans: (e)

7. The name of the following alkane is: C₃H₈

- | | |
|------------|-------------|
| (a) Ethane | (d) Pentane |
| (b) Octane | |
| (c) Hexane | (e) Propane |

Ans: (e)

8. The name of the following alkane is: C₈H₁₈

- | | |
|------------|-------------|
| (a) Ethane | (d) Pentane |
| (b) Octane | |
| (c) Hexane | (e) Propane |

Ans: (b)

9. The name of the following alkane is: C₅H₁₂

- | | |
|------------|-------------|
| (a) Ethane | (d) Pentane |
| (b) Octane | |
| (c) Hexane | (e) Propane |

Ans: (d)

Ans: (d)

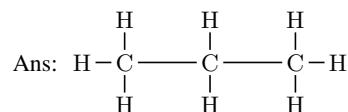
10. Write down the condensed formula for: Hexane.

Ans: CH₃ – CH₂ – CH₂ – CH₂ – CH₂ – CH₃

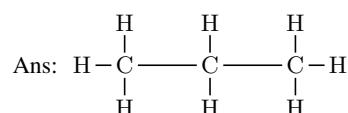
11. Write down the condensed formula for: Propane.

Ans: CH₃ – CH₂ – CH₃

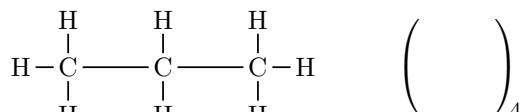
12. Write down the expanded formula for: Pentane.



13. Write down the expanded formula for: Decane.



14. Write down the molecular formula for:



Ans: C₆H₁₄, hexane

15. Write down the molecular formula for:



Ans: C₈H₁₈, octane

16. Write down the molecular formula for:



Ans: C₁₀H₂₂, decane

17. Write down the expanded formula for: hexane.



CYCLOALKANES

18. The molecular formula of cyclopropane is

- | | |
|-----------------------------------|-------------------------------------|
| (a) C ₂ H ₄ | (d) C ₆ H ₁₂ |
| (b) C ₃ H ₆ | |
| (c) C ₄ H ₈ | (e) C ₁₀ H ₂₀ |

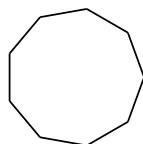
Ans: (c)

19. The molecular formula of cyclohexane is

- | | |
|-----------------------------------|-------------------------------------|
| (a) C ₂ H ₄ | (d) C ₆ H ₁₂ |
| (b) C ₃ H ₆ | |
| (c) C ₄ H ₈ | (e) C ₁₀ H ₂₀ |

Ans: (d)

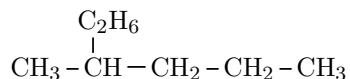
20. Name the following cycloalkane:



Ans: Cyclononane

Ans: 2,3-dimethylhexane

28. Give the name for the following compound:



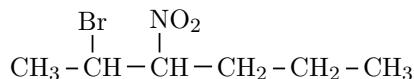
Ans: 2-ethylpentane

21. Name the following cycloalkane:



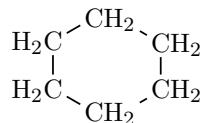
Ans: Cycloheptane

29. Give the name for the following compound:



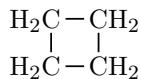
Ans: 2-bromo-3-nitrohexane

22. Name the following cycloalkane:



Ans: Cyclohexane

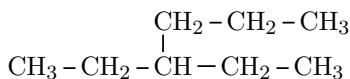
23. Name the following cycloalkane:



Ans: Cyclopropane

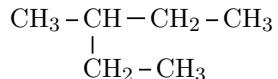
Ans: 3-ethylhexane

30. Give the name for the following compound:



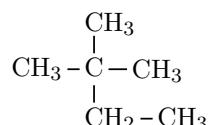
Ans: 3-methylpentane

31. Give the name for the following compound:



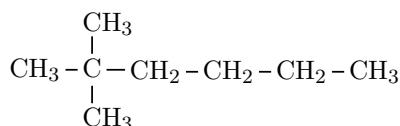
Ans: 2,3-dimethylbutane

32. Give the name for the following compound:



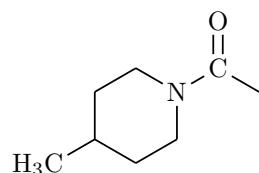
Ans: 2,3-dimethylbutane

25. Give the name for the following compound:



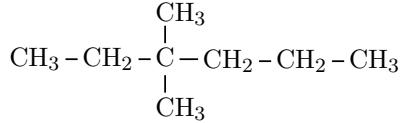
Ans: 2,2-dimethylhexane

33. Identify the functional groups in the following molecule:



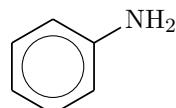
Ans: amide.

26. Give the name for the following compound:



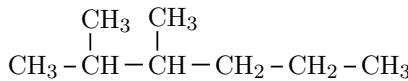
Ans: 3,3-dimethylhexane

34. Identify the functional groups in the following molecule:

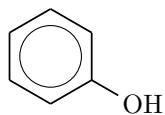


Ans: amine and aromatic.

27. Give the name for the following compound:

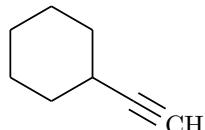


35. Identify the functional groups in the following molecule:



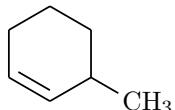
Ans: alcohol and aromatic.

36. Identify the functional groups in the following molecule:



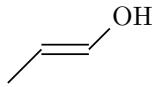
Ans: alkyne.

37. Identify the functional groups in the following molecule:



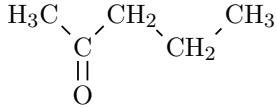
Ans: alkene.

38. Identify the functional groups in the following molecule:



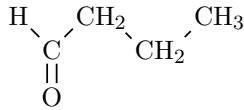
Ans: alkene and alcohol.

39. Identify the functional groups in the following molecule:



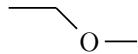
Ans: ketone.

40. Identify the functional groups in the following molecule:



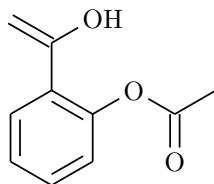
Ans: aldehyde.

41. Identify the functional groups in the following molecule:



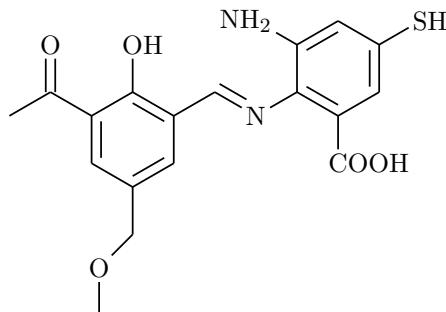
Ans: ether.

42. Identify the functional groups in the following molecule:



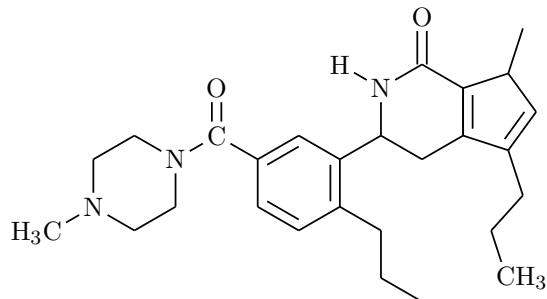
Ans: alcohol, aromatic, alkene, ester.

43. Identify the functional groups in the following molecule:



Ans: (top left) ketone; (bottom) ether; (top left) alcohol; (top right) amino and thiol; (bottom left) acid; there are also two aromatic rings

44. Identify the functional groups in the following molecule:



Ans: there are an amine, two amides, an aromatic ring and two alkanes.