

Samples	Boyle's Law
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1. A sample of chlorine gas occupies a volume of 1168 mL at a pressure of 686 mmHg. Calculate the pressure of the gas if the volume is reduced to 976 mL (assume temperature remains constant).
2. A sample of chlorine gas occupies a volume of 859 mL at a pressure of 109.98 kPa. Calculate the pressure of the gas if the volume is reduced to 133 mL (assume temperature remains constant).
3. A sample of chlorine gas occupies a volume of 1000 mL at a pressure of 691 mmHg. Calculate the pressure of the gas if the volume is reduced to 103 mL (assume temperature remains constant).