

## Yakeen NEET 2.0 2026

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## Electrochemistry

DPP: 5

- Q1** The desired amount of charge for obtaining one mole of Al from  $Al^{+3}$   
 (A)  $3 \times 96500C$   
 (B)  $96500C$   
 (C)  $\frac{96500}{3}C$   
 (D)  $\frac{96500}{2}C$
- Q2** When electricity is passed through the solution of  $AlCl_3$ , 13.5 g of Al are deposited. The number of Faraday must be  
 (A) 0.50 (B) 1.00  
 (C) 1.50 (D) 2.00
- Q3** One Faraday of electricity when passed through a solution of copper sulphate deposits  
 (A) 1 mole of Cu  
 (B) 1gm atom of Cu  
 (C) 1 molecule of Cu  
 (D) 1gm equiv. of Cu
- Q4** The density of Cu is  $8.94 \text{ g/cm}^3$  the quantity of electricity needed to plate an area  $10 \text{ cm} \times 10 \text{ cm}$  to a thickness of  $10^{-2} \text{ cm}$  using  $CuSO_4$  solution would be  
 (A) 13586C  
 (B) 27172C  
 (C) 40758C  
 (D) 20348C
- Q5** What current strength in ampere will be required to liberate 10 g of chlorine from sodium chloride solution in one hour?  
 (A) 7.55 ampere (B) 7.15 ampere  
 (C) 7.50 ampere (D) 7.45 ampere
- Q6** A certain current liberated 0.50gm of hydrogen in 2 hours. How many grams of copper can be liberated by the same current flowing for the same time in a copper sulphate solution.  
 (A) 12.7gm  
 (B) 15.9gm  
 (C) 31.8gm  
 (D) 63.5gm
- Q7** One Faraday of electricity will liberate one mole of metal from solution of  
 (A)  $AuCl_3$   
 (B)  $CuSO_4$   
 (C)  $BaCl_2$   
 (D) KCl
- Q8** One-gram metal  $M^{+2}$  was discharged by the passage of  $1.81 \times 10^{22}$  electrons. What is the atomic weight of metal?  
 (A) 33.35 (B) 133.4  
 (C) 66.7 (D) 55
- Q9** 3 Faradays of electricity was passed an aqueous solution of iron (II) bromide. The mass of iron metal (atomic mass 56) deposited at the cathode is  
 (A) 56 g  
 (B) 84 g  
 (C) 112 g



(D) 168 g

**Q10** A current of 9.65 ampere is passed through the aqueous solution of NaCl using suitable electrodes for 1000 sec. The amount of NaOH formed during electrolysis is

- (A) 2.0 g
- (B) 4.0 g
- (C) 6.0 g
- (D) 8.0 g

**Q11** A current of 2.6 ampere is passed through  $\text{CuSO}_4$  solution for 6 minute 20 seconds. The amount of Cu deposited is (At wt. of Cu = 63.5, Faraday = 96500C )

- (A) 6.35 g
- (B) 0.635 g
- (C) 0.325 g
- (D) 3.175 g

**Q12** 3 Faradays of electricity are passed through molten  $\text{Al}_2\text{O}_3$ , aqueous solution of  $\text{CuSO}_4$  and molten NaCl taken in three different electrolytic cells. The amount of Al, Cu and Na deposited at the cathodes will be in the ratio

- (A) 1 mole : 2 mole : 3 mole
- (B) 1 mole : 1.5 mole : 3 mole
- (C) 3 mole : 2 mole : 1 mole
- (D) 1 mole : 1.5 mole : 2 mole



## Answer Key

Q1 (A)

Q2 (C)

Q3 (D)

Q4 (B)

Q5 (A)

Q6 (B)

Q7 (D)

Q8 (C)

Q9 (B)

Q10 (B)

Q11 (C)

Q12 (B)



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