

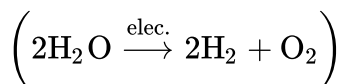
Yakeen NEET 2.0 2026

Physical Chemistry By Amit Mahajan Sir

Some Basic Concept of Chemistry

DPP: 5

- Q1** The number of moles of oxygen obtained by the electrolytic decomposition of 90 g water is

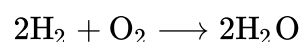


- (A) 2.5 (B) 5
(C) 7.5 (D) 10
- Q2** The number of moles of hydrogen molecules required to produce 20 moles of ammonia through Habers process is .
(A) 40 (B) 10
(C) 20 (D) 30
- Q3** The volume of a gas at 0°C and 760 mm pressure is 22.4 cc. The no. of molecules present in this volume is
(A) $10^{-3} N_A$
(B) $10^{-4} N_A$
(C) $10^{-5} N_A$
(D) $10^{-2} N_A$
- Q4** 22.4 litre of water vapour at NTP, when condensed to water, occupies an approximate volume of:
(A) 18 L
(B) 1 L
(C) 1 mL
(D) 18 mL
- Q5** From 320mg. of O_2 , 6.023×10^{20} molecules are removed, the no. of moles remained are:
(A) 9×10^{-3} moles
(B) 9×10^{-2} moles
(C) Zero
(D) 3×10^{-3} moles
- Q6** The number of moles of hydrogen molecules required to produce 40 moles of ammonia

through Haber's process is:

- (A) 60 (B) 30
(C) 40 (D) 50

- Q7** Calculate the moles of H_2O vapours formed if 1.57 mole of O_2 are used in presence of excess of H_2 for the given change.



- Q8** How many molecule are present in 12 L of liquid CCl_4 ? The density of the liquid is 1.59 g cm^{-3} ?
(A) 7.44×10^{26}
(B) 1.59×10^{26}
(C) 0.744×10^{26}
(D) 15.9×10^{26}
- Q9** How many grams of calcium oxide is obtained on heating 100 g of CaCO_3 (s) ?
(A) 50 g
(B) 40 g
(C) 56 g
(D) 44 g
- Q10** What is the mass of glucose required to produce 44 g of CO_2 on complete combustion?
(A) 30 g
(B) 45 g
(C) 60 g
(D) 22 g
- Q11** According to the following reaction the minimum quantity in g of H_2S needed to precipitate 63.5gm of Cu^{2+} ions will be nearly?
 $\text{Cu}^{+2} + \text{H}_2\text{S} \rightarrow \text{CuS} + 2\text{H}^+$
(A) 63.5 g
(B) 31.75g
(C) 34g



(D) 20g

Q12 What is the weight of oxygen required for the complete combustion of 2.8 kg of ethylene?

- (A) 2.8 kg
- (B) 6.4 kg
- (C) 9.6 kg
- (D) 96 kg



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Answer Key

Q1 (A)

Q2 (D)

Q3 (A)

Q4 (D)

Q5 (A)

Q6 (A)

Q7 3.14 moles H_2O

Q8 (C)

Q9 (C)

Q10 (A)

Q11 (C)

Q12 (C)

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