

Yakeen NEET 2.0 2026

Physical Chemistry By Amit Mahajan Sir

Some Basic Concept of Chemistry

DPP: 6

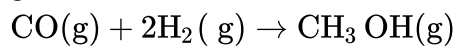
- Q1** The number of moles of 'B' required to produce 2.5 mole of AB_4 according to equation:
 $A + 4B \rightarrow AB_4$
 (A) 1 (B) 10
 (C) 5 (D) 15
- Q2** The equation $2Al(s) + \frac{3}{2}O_2 \rightarrow Al_2O_3(s)$ shows that
 (A) 2 moles of aluminium react with $\frac{3}{2}$ moles of oxygen to produce one mole of aluminium oxide
 (B) 2 atoms of aluminium react with $\frac{3}{2}$ atoms of oxygen to produce one atom of aluminium oxide
 (C) 2g of aluminium react with $\frac{3}{2}$ g of oxygen to produce 1 g of aluminium oxide
 (D) 2 g of aluminium react with $\frac{3}{2}$ litres of oxygen to produce 1 g of aluminium oxide
- Q3** The number of moles of $Ca(HCO_3)_2$ required to prepare 1.50 moles of CO_2 , according to the equation
 $Ca(HCO_3)_2 + 2HCl \rightarrow CaCl_2 + 2CO_2 + 2H_2O$
 (A) 1.00 mol
 (B) 2.00 mol
 (C) 0.750 mol
 (D) 1.50 mol
- Q4** 0.7 moles of potassium sulphate is allowed to react with 0.9 moles of barium chloride in aqueous solutions. The number of moles of the substance precipitated in the reaction is:
 (A) 1.4 moles of potassium chloride
 (B) 0.7 moles of barium sulphate
 (C) 1.6 moles of potassium chloride
 (D) 1.6 moles of barium sulphate
- Q5** How many grams of calcium oxide is obtained on heating 100 g of $CaCO_3(s)$?
 (A) 50 g
 (B) 40 g
 (C) 56 g
 (D) 44 g
- Q6** What is the mass of glucose required to produce 44 g of CO_2 on complete combustion?
 (A) 30 g
 (B) 45 g
 (C) 60 g
 (D) 22 g
- Q7** What is the weight of oxygen required for the complete combustion of 2.8 kg of ethylene?
 (A) 2.8 kg
 (B) 6.4 kg
 (C) 9.6 kg
 (D) 96 kg
- Q8** According to the following reaction the minimum quantity in g of H_2S needed to precipitate 63.5gm of Cu^{2+} ions will be nearly?
 $Cu^{+2} + H_2S \rightarrow CuS + 2H^+$
 (A) 63.5 g
 (B) 31.75g
 (C) 34g
 (D) 20g
- Q9** A quantity of 2.76 g of silver carbonate on being strongly heated yields a residue(Solid/liquid) weighing *Gramatomicmass of* ($A_g = 108$), $C=12g$ and $O=16g$)
 $Ag_2CO_3(s) \xrightarrow{\Delta} 2Ag(s) + CO_2(g) + \frac{1}{2}O_2(g)$
 (A) 2.16 g



- (B) 2.48 g
 (C) 2.32 g
 (D) 2.64 g

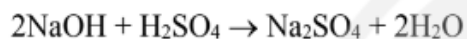
Q10 How many grams of H_2 are required to consume 2 mol of CO ?

(Gram atomic mass of H = 1 g, C = 12 g and O = 16 g)



- (A) 2 g
 (B) 4 g
 (C) 8g
 (D) 16 g

Q11 The mass of 70% H_2SO_4 required for neutralization of one mole of NaOH is : (Gram atomic mass of H = 1 g, O = 16 g, Na = 23 g and S = 32 g)



- (A) 70 g
 (B) 35 g
 (C) 30 g
 (D) 95 g



Answer Key

Q1 (B)

Q2 (A)

Q3 (C)

Q4 (B)

Q5 (C)

Q6 (A)

Q7 (C)

Q8 (C)

Q9 (A)

Q10 (C)

Q11 (A)



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