

## Yakeen NEET 2.0 2026

## Physical Chemistry By Amit Mahajan Sir

## Some Basic Concept of Chemistry

DPP: 9

**Q1**  $B_1$  g of an element gives  $B_2$  g of its chloride, the equivalent mass of the element is:

- (A)  $\frac{B_1}{B_2 - B_1} \times 35.5$   
 (B)  $\frac{B_2}{B_2 - B_1} \times 35.5$   
 (C)  $\frac{B_2 - B_1}{B_1} \times 35.5$   
 (D)  $\frac{B_2 - B_1}{B_2} \times 35.5$

**Q2** Consider the following statements.

Statement-I: Equivalent weight of ozone in the change  $O_3 \rightarrow O_2$  is 8.

Statement-II: 1 mole of  $O_3$  on decomposition gives  $\frac{3}{2}$  moles of  $O_2$ .

Choose the correct option.

- (A) Both Statement-I and Statement-II are correct.  
 (B) Both Statement-I and Statement-II are incorrect.  
 (C) Statement-I is correct and Statement-II is incorrect.  
 (D) Statement-I is incorrect and Statement-II is correct.

**Q3** An oxide of a metal M has 60% metal by mass. What is the equivalent weight of the metal?

- (A) 12 (B) 24  
 (C) 36 (D) 48

**Q4** If  $m_1$  g of a metal A, displaces  $m_2$  g of another metal B from its salt solution and if their equivalent masses are  $E_1$  and  $E_2$  respectively, then the equivalent mass of A can be expressed as  
 (A)

$$E_1 = \frac{m_2 \times E_2}{m_1}$$

$$(B) E_1 = \frac{m_1}{m_2} \times E_2$$

$$(C) E_1 = \frac{m_1 \times m_2}{E_2}$$

$$(D) E_1 = \sqrt{\frac{m_1}{m_2}} \times E_2$$

**Q5** If the weight of metal oxide is  $x$  g containing  $y$  g of oxygen, the equivalent weight of metal will be

- (A)  $E = \frac{8x}{y}$  (B)  $E = \frac{8(y-x)}{x}$   
 (C)  $E = \frac{y}{8}$  (D)  $E = \frac{8(x-y)}{y}$

**Q6** 2 g of metal on reaction with  $H_2SO_4$  give 5 g metal sulphate. Find equivalent mass of metal?

- (A) 2 (B) 16  
 (C) 32 (D) 64

**Q7** Carbon forms two oxides which have different compositions. The equivalent mass of which remains constant?

- (A) carbon  
 (B) oxygen  
 (C) neither carbon nor oxygen  
 (D) both carbon and oxygen

**Q8** 4 g of a metal oxide contains 1.6 g oxygen, then equivalent mass of the metal is

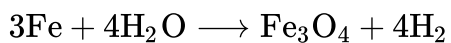
- (A) 3.2 (B) 24  
 (C) 12 (D) 20

**Q9** When a metal is burnt, its mass is increased by 24 per cent. The equivalent mass of the metal will be:

- (A) 25 (B) 24  
 (C) 33.3 (D) 76



**Q10** In the following change,



If the atomic mass of iron is 56, then its equivalent mass will be:

- (A) 42 (B) 63  
(C) 21 (D) 84

**Q11** Sulphur forms two chlorides  $\text{S}_2\text{Cl}_2$  and  $\text{SCl}_2$ .

The equivalent mass of sulphur in  $\text{SCl}_2$  is 16.

The equivalent weight of sulphur in  $\text{S}_2\text{Cl}_2$  is

- (A) 8 (B) 16  
(C) 32 (D) 64

**Q12** The oxide of a metal has 32% oxygen. Its equivalent weight would be

- (A) 34 (B) 32  
(C) 17 (D) 16



## Answer Key

Q1 (A)

Q2 (A)

Q3 (A)

Q4 (B)

Q5 (D)

Q6 (C)

Q7 (B)

Q8 (C)

Q9 (C)

Q10 (C)

Q11 (C)

Q12 (C)



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