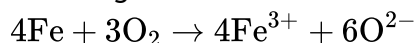


Yakeen NEET 2.0 2026

Physical Chemistry By Amit Mahajan Sir

DPP: 2

Redox Reaction

Q1 Following reaction describes the rusting of iron

Which one of the following statement is incorrect?

- (A) This is an example of a redox reaction
- (B) Metallic iron is reduced to Fe^{3+}
- (C) Fe^{3+} is an oxidizing agent
- (D) Metallic iron is a reducing agent

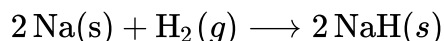
Q2 ASSERTION & REASON

Assertion (A): A substance which gets reduced can act as an oxidising agent.

Reason (R): In the reaction, $3\text{ClO}^- \rightarrow \text{ClO}_3^- + 2\text{Cl}^-$, Cl atom is oxidised as well as reduced.

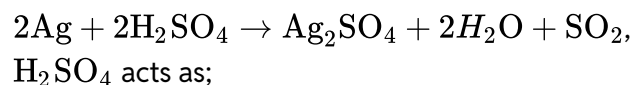
Choose the correct option.

- (A) Both Assertion (A) and Reason (R) are True and the Reason (R) is a correct explanation of the Assertion (A).
- (B) Both Assertion (A) and Reason (R) are True but Reason (R) is not a correct explanation of the Assertion (A).
- (C) Assertion (A) is True but the Reason (R) is False.
- (D) Assertion (A) is False but Reason (R) is True.

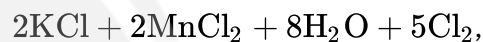
Q3 In the reaction given below, identify the species undergoing redox reaction

- (A) Na is reduced and hydrogen is oxidised
- (B) Na is oxidised and hydrogen is reduced
- (C) Na undergoes oxidation and hydrogen undergoes reduction

(D) Both B and C

Q4 In a reaction

- (A) Reducing agent
- (B) Oxidising agent
- (C) Dehydrate agent
- (D) None of these

Q5 In the reaction, $2\text{KMnO}_4 + 16\text{HCl} \rightarrow$ 

the reduction product is :

- (A) Cl_2
- (B) KCl
- (C) MnCl_2
- (D) H_2O

Q6 For the unbalanced reaction

$\text{AX} + \text{BY} + \text{H}_2\text{O} \rightarrow \text{HA} + \text{OY} + \text{X}_2\text{B}$. Let the oxidation number of X be -2 and X, H_2O are not involved in redox reaction. The element(s) undergoing oxidation is:

- (A) A
- (B) B
- (C) Y
- (D) Both B and Y

Q7 **Assertion (A):** A substance which gets reduced can act as an oxidising agent.

Reason (R): In the reaction, $3\text{ClO}^- \rightarrow \text{ClO}_3^- + 2\text{Cl}^-$, Cl atom is oxidised as well as reduced.

- (A) Both Assertion (A) and Reason (R) are True and the Reason (R) is a correct explanation of the Assertion (A).



- (B) Both Assertion (A) and Reason (R) are True
but Reason (R) is not a correct explanation of
the Assertion (A).
- (C) Assertion (A) is True but the Reason (R) is
False.
- (D) Assertion (A) is False but Reason (R) is True.

Q8 Which reaction does **not** involve neither
oxidation nor reduction?

- (A) $\text{VO}^{2+} \rightarrow \text{V}_2\text{O}_3$
- (B) $\text{Na} \rightarrow \text{Na}^+$
- (C) $\text{CrO}_4^{2-} \rightarrow \text{Cr}_2\text{O}_7^{2-}$
- (D) $\text{Zn}^{2+} \rightarrow \text{Zn}$

Q9 In which of the following reactions, there is no
change in valency?

- (A) $4\text{KClO}_3 \rightarrow 3\text{KClO}_4 + \text{KCl}$
- (B) $\text{SO}_2 + 2\text{H}_2\text{S} \rightarrow 2\text{H}_2\text{O} + 3\text{S}$
- (C) $\text{BaO}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + \text{H}_2\text{O}_2$
- (D) $3\text{BaO} + \text{O}_2 \rightarrow 2\text{BaO}_2$

Q10 Which of the following is not an example of
redox reaction?

- (A) $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$
- (B) $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
- (C) $2\text{K} + \text{F}_2 \rightarrow 2\text{KF}$
- (D) $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{HCl}$



Answer Key

Q1 (B)

Q2 (B)

Q3 (D)

Q4 (B)

Q5 (C)

Q6 (D)

Q7 (B)

Q8 (C)

Q9 (C)

Q10 (D)



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