

Yakeen NEET 2.0 2026

Physical Chemistry

Electrochemistry

DPP: 1

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Q1 The standard reduction potential values of three metallic cations X, Y and Z are 0.52, -3.03 and -1.18 V respectively. The order of reducing power of the corresponding metal is

- (A) $Y > Z > X$
 (B) $X > Y > Z$
 (C) $Z > Y > X$
 (D) $Z > X > Y$

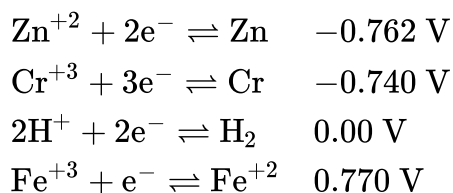
Q2 To a mixture containing pieces of Zn, Cu and silver, $1\text{M H}_2\text{SO}_4$ was added. H_2 gas was found to be evolved. Which of the metal/metals do you think has/have reacted?

$$E_{\text{Zn}^{+2}/\text{Zn}}^{\circ} = -0.76 \text{ V} \quad E_{\text{Cu}^{+2}/\text{Cu}}^{\circ} = 0.34 \text{ V}$$

$$E_{\text{Ag}^{+2}/\text{Ag}}^{\circ} = 0.80 \text{ V}$$

- (A) All the metals
 (B) Only Zn
 (C) Both Zn and Cu
 (D) Only Ag

Q3 The standard reduction potentials at 298 K for the following half reactions are given



Which is the strongest reducing agent?

- (A) Zn

- (B) Cr
 (C) H_2
 (D) Fe^{+2}

Q4 When Zn dust is added to a solution of MgCl_2

- (A) No reaction will take place
 (B) ZnCl_2 is formed
 (C) Zinc dissolved in the solution
 (D) Magnesium is precipitated

Q5 The standard reduction potential of A, B and C are 0.34 V, 0.80 V and 0.79 V respectively. The decreasing order of deposition of metals on electrodes are

- (A) $A > B > C$
 (B) $B > C > A$
 (C) $C > B > A$
 (D) $A > C > B$

Q6 Using the data given, find strongest oxidizing agent.

$$E_{\text{Cl}_2/\text{Cl}^-}^{\circ} = 1.36 \text{ V}$$

$$E_{\text{Cr}^{+6}/\text{Cr}^{+3}}^{\circ} = 1.33 \text{ V}$$

$$E_{\text{MnO}_4^-/\text{Mn}^{+2}}^{\circ} = 1.51 \text{ V}$$

$$E_{\text{Cr}^{+3}/\text{Cr}}^{\circ} = -0.74 \text{ V}$$

- (A) Cl^-
 (B) Cr
 (C) Cr^{+3}
 (D) MnO_4^-

Q7



A metal having negative reduction potential when dipped in the solution its own ions, has a tendency

- (A) to pass into the solution
- (B) to be deposited from the solution
- (C) to become electrically positive
- (D) to remain neutral

Q8 Zn cannot displace the following ion from its aqueous solution:

- (A) Ag^+
- (B) Cu^{+2}
- (C) Fe^{+2}
- (D) Na^+

Q9 Which of the following displacement does not occur?

- (A) $\text{Zn} + 2\text{H}^+ \rightarrow \text{Zn}^{+2} + \text{H}_2$
- (B) $\text{Fe} + 2\text{Ag}^+ \rightarrow \text{Fe}^{+2} + 2\text{Ag}$
- (C) $\text{Cu} + \text{Fe}^{+2} \rightarrow \text{Fe} + \text{Cu}^{+2}$
- (D) $\text{Zn} + \text{Pb}^{+2} \rightarrow \text{Zn}^{+2} + \text{Pb}$

Q10 The Kohlrausch's law is related to

- (A) Conductance of ions at infinite dilution.
- (B) Independent migration of ions.
- (C) Both (A) & (B)
- (D) Neither (A) & (B)



Answer Key

Q1 (A)

Q2 (B)

Q3 (A)

Q4 (A)

Q5 (B)

Q6 (D)

Q7 (A)

Q8 (D)

Q9 (C)

Q10 (C)



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