Magarmach Practice Questions (MPQ)

What is the unit of absolute atomic mass?

(A) amu

(B) grams

(C) kg

(D) g/mol

$$_4\text{He}$$

R.A.M. = 4

A.A.M. = 4 amu

G.A.M. = 4g

P
W

Relative atomic mass is the ratio of the average mass of atoms of an element to:

$$R.A.M = \frac{\text{Mass of atom}}{1 \text{ a.m.u. or } \frac{1}{12} \text{ of mass of } C-12 \text{ atom}}$$

(A) 1 amu

(B) 1 gram

(C) $1/12^{th}$ of mass of C-12 atom

(D) 1 mole

What is the relative atomic mass of oxygen?

A) 16

B) 32

C) 8

D) 12

What is the absolute atomic mass of hydrogen approximately?

(A) 1 g

(B) 1.67×10^{-24} g

(C) 1.67×10^{-27} kg

(D) Both b and c

Q_H A.A.M. = 1 a.m.u.
$$= 1.67 \times 10^{-24} \text{ g}$$
$$= 1.67 \times 10^{-27} \text{ kg}$$

The gram atomic mass of nitrogen is : $_7^{14}N$

A 14 g

B 7 g

C 28 g

D 1 g

Which of the following is true for relative atomic mass?

(A) It has units

(B) It is a ratio and has no units

(C) Measured in grams

(D) Measured in kilograms

Which is the correct value for Avogadro's number?

A 6.022×10^{22}

B 6.022×10^{24}

C 6.022×10^{23}

D 3.011×10^{23}

The gram atomic mass of an element is numerically equal to its:

$$\binom{14}{7}N$$

$$14g = G.A.M. \text{ mass of } \underline{N_A} \text{ atoms}$$

(A) ✗ Absolute mass

(B) Molecular mass

(C) ✗ Relative atomic mass

(D) ✓ Molar mass

The absolute atomic mass of carbon is approximately:

$$^{12}_{6}C \qquad 12u$$

A 12 g

B 1.99×10^{-23} g

C 1.99×10^{-22} g

D 12 amu

What is the gram atomic mass of sulfur?

$$\overset{32}{\underset{16}{S}}$$

A) 16 g

B) 32 g

C) 64 g

D) 12 g

THANK YOU