

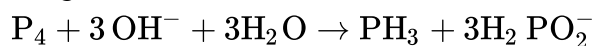
Yakeen NEET 2.0 2026

Physical Chemistry By Amit Mahajan Sir

DPP: 3

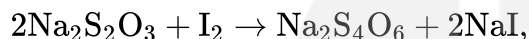
Redox Reaction

Q1 Identify the **correct** statements with reference to the given reaction



- (A) Phosphorus is undergoing reduction only.
 (B) Phosphorus is undergoing oxidation only.
 (C) Phosphorus is undergoing oxidation as well as reduction.
 (D) Phosphorus is undergoing neither oxidation nor reduction.

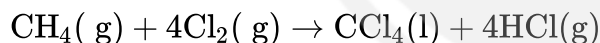
Q2 In the reaction,



the oxidation state of sulphur is:

- (A) decreased (B) increased
 (C) unchanged (D) None of these

Q3 What is the change in oxidation number of carbon in the following reaction?

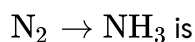


- (A) 0 to +4 (B) -4 to +4
 (C) 0 to -4 (D) +4 to +4

Q4 The oxidation state of *Ni* in $Ni_{0.98}O_{1.00}$ is:

- (A) $-\frac{49}{100}$
 (B) $+\frac{100}{49}$
 (C) $-\frac{100}{49}$
 (D) $+\frac{49}{100}$

Q5 Equivalent weight of N_2 in the change

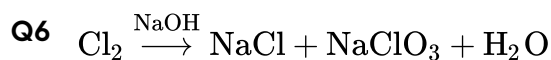


- (A) $\frac{28}{6}$

(B) 28

(C) $\frac{28}{2}$

(D) $\frac{28}{3}$



The equivalent mass of Cl_2 in the above reaction is

- (A) *M*
 (B) $M/3$
 (C) $M/2$
 (D) $3M/5$

Q7 The equivalent weight of FeS_2 in the following reaction is $FeS_2 + O_2 \rightarrow Fe^{3+} + SO_2$

- (A) $\frac{Mol.wt}{1}$
 (B) $\frac{Mol.wt}{7}$
 (C) $\frac{Mol.wt}{11}$
 (D) $\frac{Mol.wt}{9}$

Q8 Equivalent weight of FeC_2O_4 in the change $FeC_2O_4 \rightarrow Fe^{3+} + CO_2$ is

- (A) $M/3$
 (B) $M/6$
 (C) $M/2$
 (D) $M/1$

Q9 In the following change, $3Fe + 4H_2O \rightarrow Fe_3O_4 + 4H_2$. If the atomic weight of iron is 56, then its equivalent weight will be

- (A) 42 (B) 21
 (C) 63 (D) 84

Q10 When HNO_3 is converted into NH_3 , the equivalent weight of HNO_3 will be



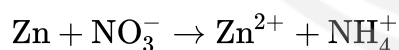
- (A) $M/2$
 (B) $M/1$
 (C) $M/6$
 (D) $M/8$

- Q11** The equivalent weight of phosphoric acid (H_3PO_4) in the reaction
 $NaOH + H_3PO_4 \rightarrow NaH_2PO_4 + H_2O$
 (A) 59 (B) 49
 (C) 25 (D) 98

- Q12** Choose the set of coefficients that **correctly** balances the following equation.
 $xCr_2O_7^{2-} + yH^+ + ze^- \rightarrow aCr^{3+} + bH_2O$
 (A) $x = 2, y = 14, z = 6, a = 2, b = 7$
 (B) $x = 1, y = 14, z = 6, a = 2, b = 7$
 (C) $x = 2, y = 7, z = 6, a = 2, b = 7$
 (D) $x = 2, y = 7, z = 6, a = 1, b = 7$

- Q13** Consider the following reaction,
 $xMnO_4^- + yC_2O_4^{2-} + zH^+ \rightarrow xMn^{2+} + 2yCO_2 + \frac{z}{2}H_2O$
 The value of x, y and z in the above reaction are respectively:
 (A) 5, 2 and 6 (B) 2, 5 and 8
 (C) 2, 5 and 16 (D) 5, 2 and 8

- Q14** For the redox reaction:



In basic medium, coefficient of Zn, NO_3^- and OH^- in the balanced equation respectively are;
 (A) 4, 1, 7 (B) 7, 4, 1
 (C) 4, 1, 10 (D) 1, 4, 10

- Q15** In the balanced chemical reaction,
 $IO_3^- + aI^- + bH^+ \rightarrow cH_2O + dI_2$
 a, b, c and d respectively correspond to
 (A) 5, 6, 3, 3 (B) 5, 3, 6, 3
 (C) 3, 5, 3, 6 (D) 5, 6, 5, 5



Answer Key

Q1 (C)

Q2 (B)

Q3 (B)

Q4 (B)

Q5 (A)

Q6 (D)

Q7 (C)

Q8 (A)

Q9 (B)

Q10 (D)

Q11 (D)

Q12 (B)

Q13 (C)

Q14 (C)

Q15 (A)



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