

Yakeen NEET 2.0 (2026)

Physical Chemistry By Amit Mahajan Sir

DPP: 4

Redox Reaction

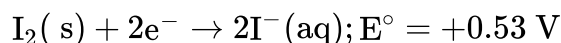
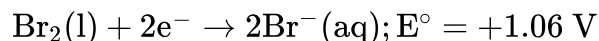
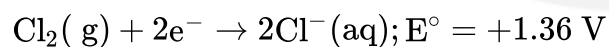
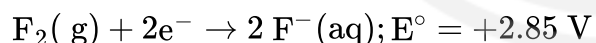
Q1 During the preparation of standard Mohr's salt solution for the titration against potassium permanganate, dilute sulphuric acid is added to

- (A) To prevent the hydrolysis of ferrous ions
- (B) To prevent the hydrolysis of ferric ions
- (C) To prevent the hydrolysis of ammonium ions.
- (D) To prevent the hydrolysis of sulphate ions.

Q2 The appearance of Mohr's salt crystals is

- (A) White
- (B) Pale yellow
- (C) Light green
- (D) Light blue

Q3 Standard reduction potentials of the half reactions are given below:



The strongest oxidizing and reducing agents respectively are

- (A) F_2 and I^-
- (B) Br_2 and Cl^-
- (C) Cl_2 and Br^-
- (D) Cl_2 and I_2

Q4 Electrode potential depends upon

- (A) Size of electrode
- (B) Surface area of electrode
- (C) Temperature
- (D) Shape of electrode

Q5 Standard reduction electrode potential of three metals X, Y and Z are -1.2 V , $+0.5 \text{ V}$ and -3.0 V respectively. The reducing power of these metals will be

- (A) $X < Y > Z$
- (B) $Y > Z > X$
- (C) $Y > X > Z$
- (D) $Z > X > Y$

Q6 In alkaline medium ClO_2 oxidises H_2O_2 to O_2 and reduced itself to Cl^- , then how many moles of H_2O_2 will be oxidised by one mole of ClO_2 ?

- (A) 1.0
- (B) 1.5
- (C) 2.5
- (D) 3.5

Q7 Equivalent weight of FeC_2O_4 in the change:



- (A) $M/3$
- (B) $M/6$
- (C) $M/2$
- (D) $M/1$

Q8 How many moles of FeSO_4 reacts with one mole of KMnO_4 in acidic medium?

- (A) $2/5$
- (B) 5
- (C) $1/2$
- (D) $1/5$

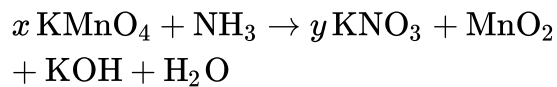
Q9 The coefficients of I^- , IO_3^- and H^+ in the redox reaction,



In the balanced form respectively are:

- (A) 5,1,6 (B) 1, 5, 6
(C) 6,1,5 (D) 5,6,1

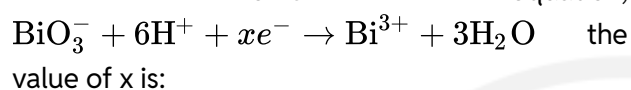
Q10 In the redox reaction,



x and y are:

- (A) x=4, y=6 (B) x=3, y=8
(C) x=8, y=6 (D) x=8, y=3

Q11 In ionic equation,



- (A) 6 (B) 2
(C) 4 (D) 3



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Answer Key

Q1 (A)

Q2 (C)

Q3 (A)

Q4 (C)

Q5 (D)

Q6 (C)

Q7 (A)

Q8 (B)

Q9 (A)

Q10 (D)

Q11 (B)



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