

YAKEEN NEET 2.0

2026

Ionic Equilibrium




Physical Chemistry

Lecture -01

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Topics *to be covered*

-  **A** Revision of Last Class
-  **B** Concepts of Acids & Bases
-  **C** Ostwald Dilution Law



Rule to Attend Class




- 1. Always sit in a peaceful environment with headphone and be ready with your copy and pen.**
- 2. Never ever attend a class from in between or don't join a live class in the middle of the chapter.**
- 3. Make sure to revise the last class before attending the next class & always complete your home work along with DPP.**
- 4. Never ever engage in chat whether live or recorded on the topic which is not being discussed in current class as by doing so u can be blocked by the admin team or your subscription can be cancelled.**



Rule to Attend Class



- 5. Try to make maximum notes during the class if something is left then u can use the notes pdf after the class to complete the remaining class.**
- 6. Always ask your doubts in doubt section to get answer from faculty. Before asking any doubt please check whether same doubt has been asked by someone or not.**
- 7. Don't watch the videos in high speed if you want to understand better.**



There is one big flaw in your Preparation that's name is Backlog ? What do we say to Backlog ?



NOT TODAY !!!

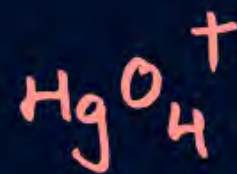


Various Concept of Acids & Bases





Arrhenius Concept of Acids & Bases



Acid ÷ acid + water \longrightarrow H^+ or H_3O^+ + anion



Base ÷ base + water \longrightarrow OH^- + Cation



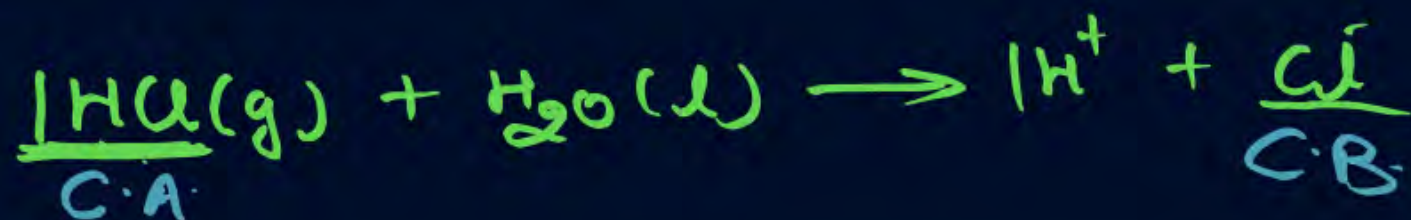


Bronsted Lowry Concept of Acids & Bases

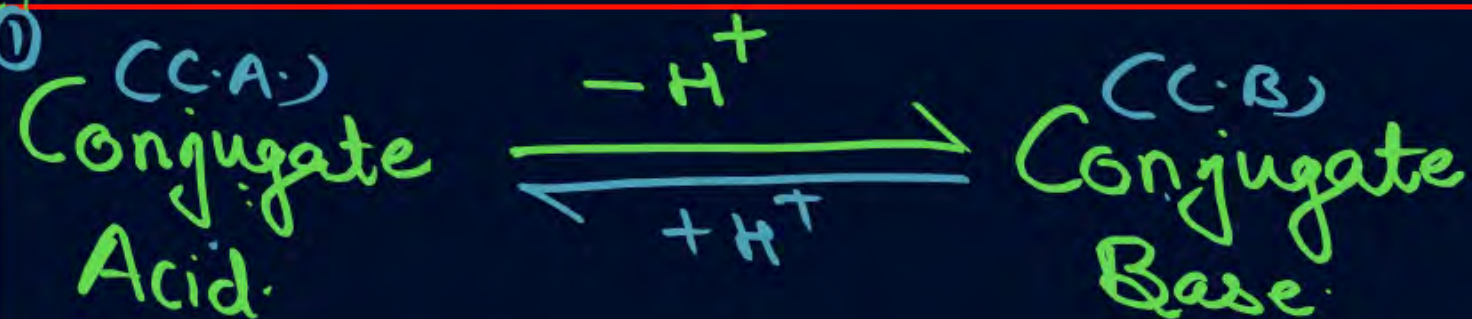


Bronsted Acid

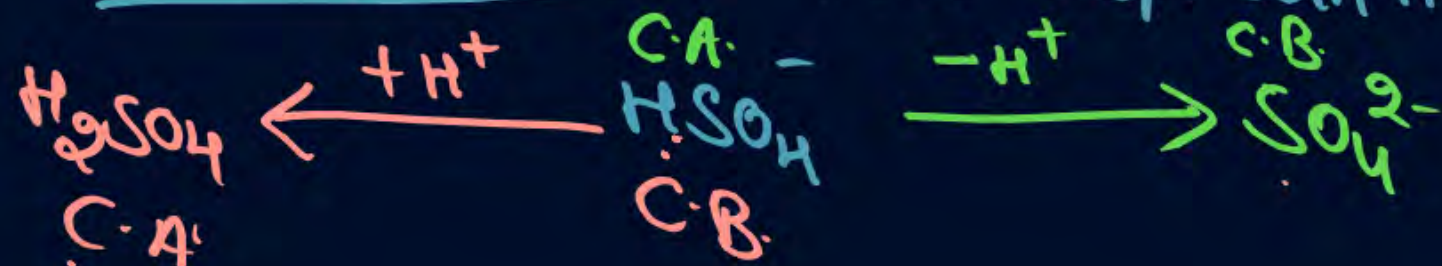
↓
Capable of donating H^+ ion or H_3O^+ ion.



M.T.T



② Amphiprotic ÷ donate & accept both H^+ ion

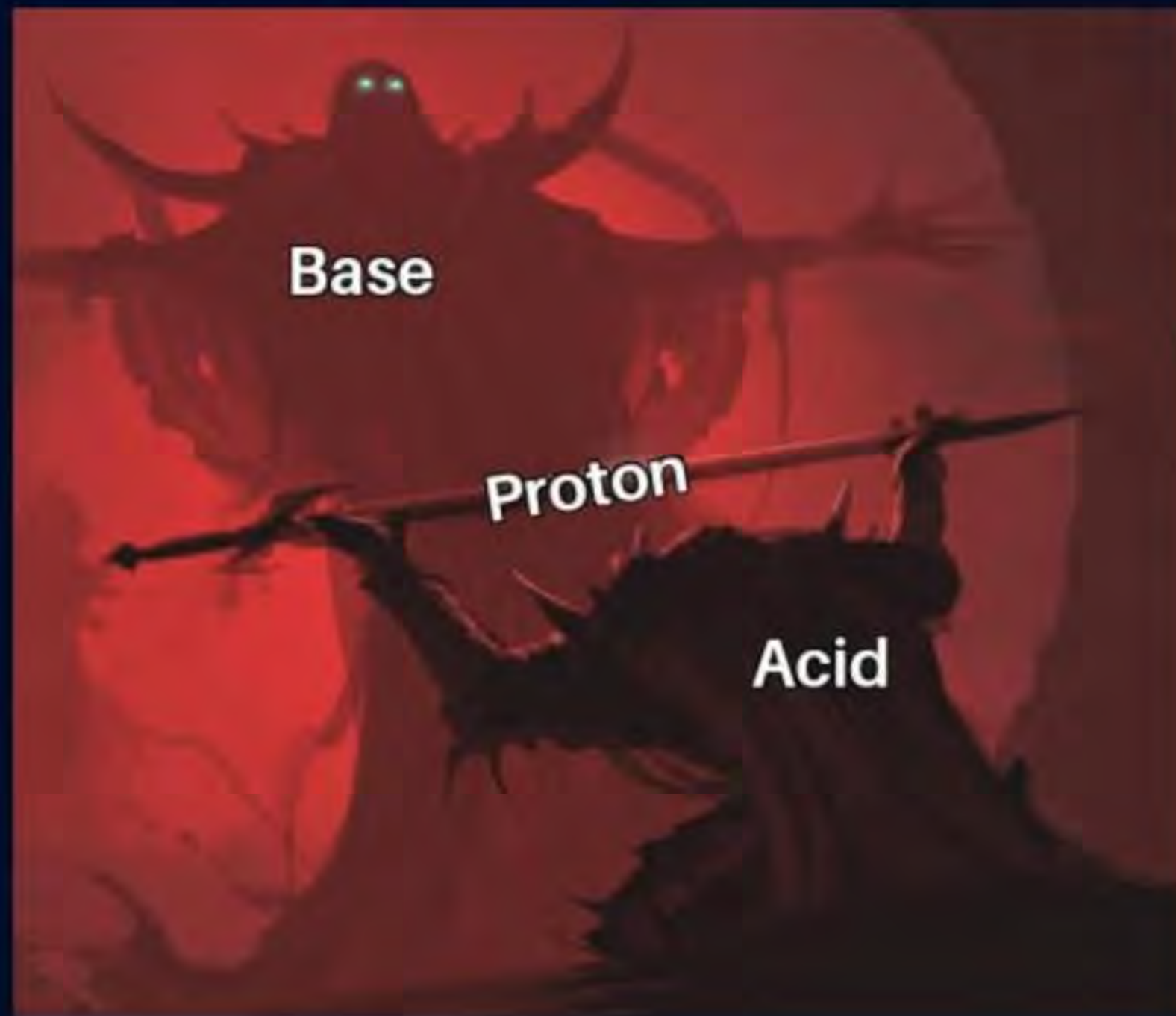
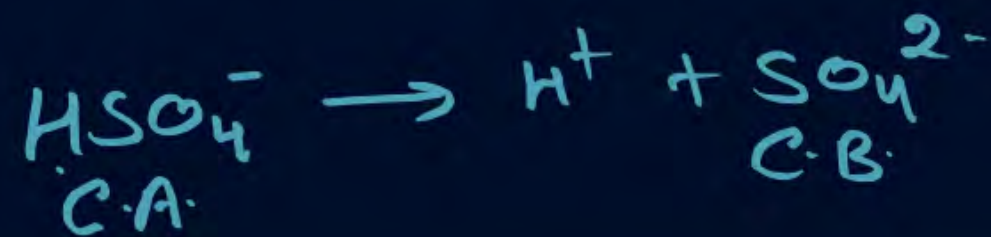
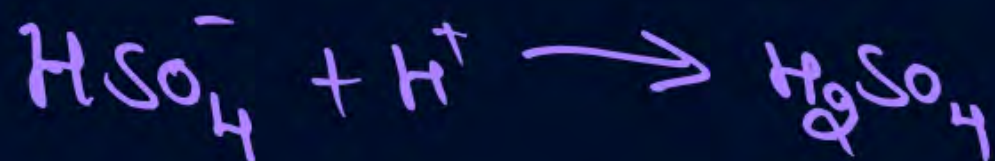
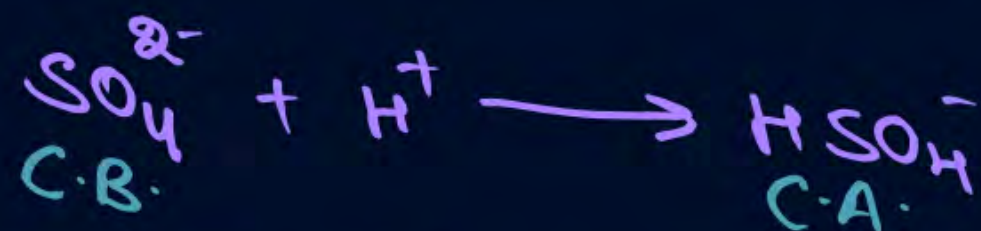
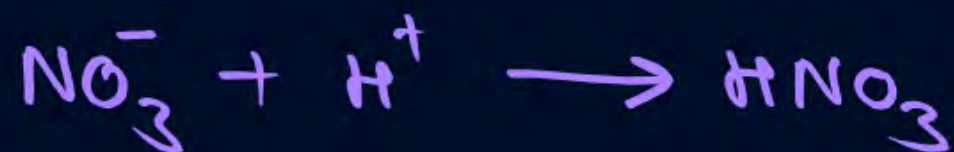
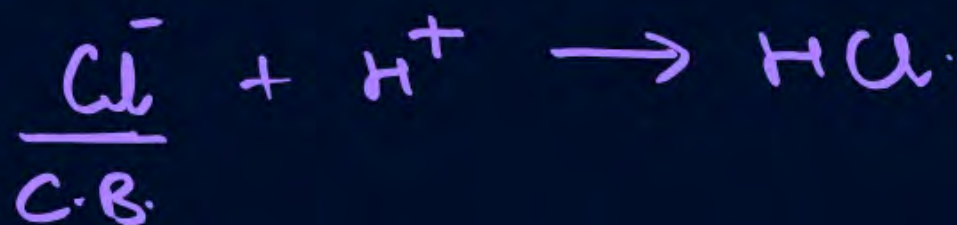




Bronsted Base

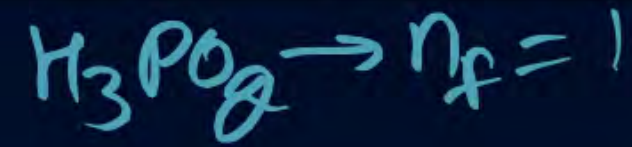
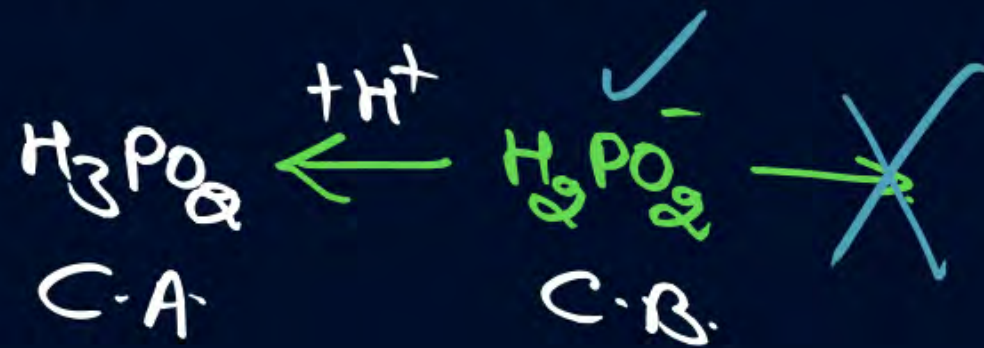
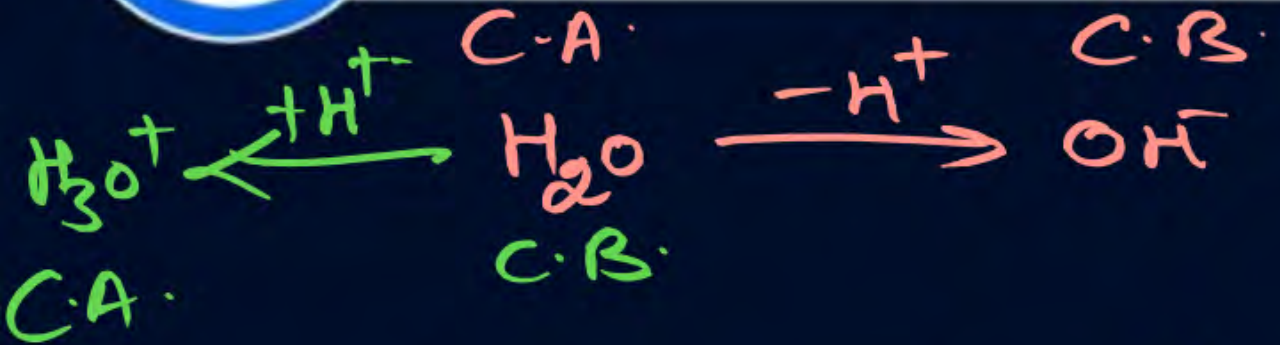


Capable of accepting H^+ (Proton)

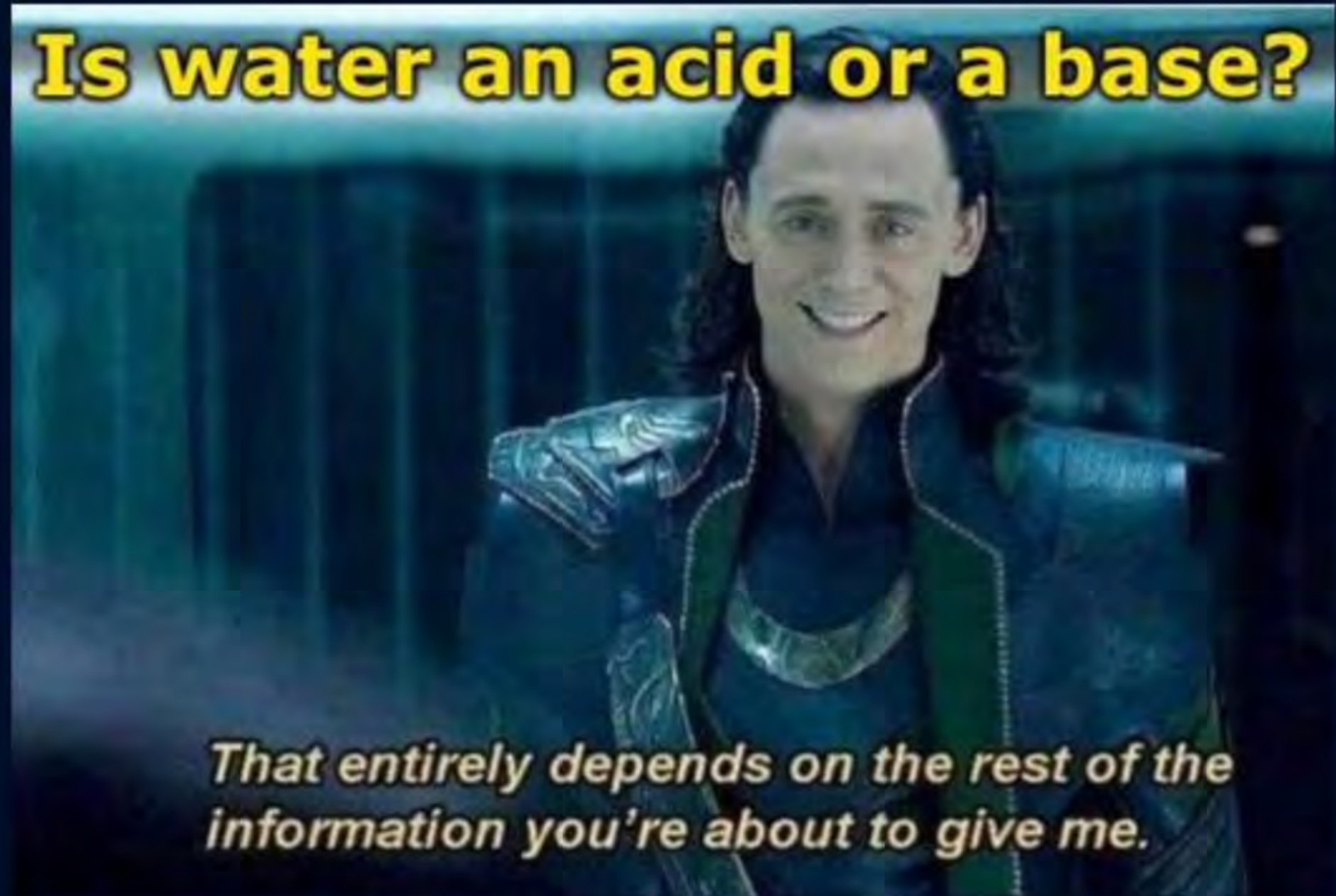




Bronsted Lowry Concept of Acids & Bases



Is water an acid or a base?

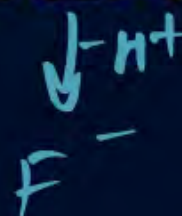
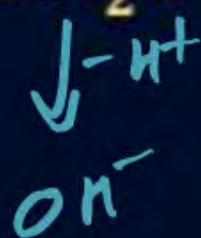


That entirely depends on the rest of the information you're about to give me.

QUESTION – (NEET 2019)

Conjugate base for Bronsted acids H_2O and HF are:

A OH^- and H_2F^+ , respectively



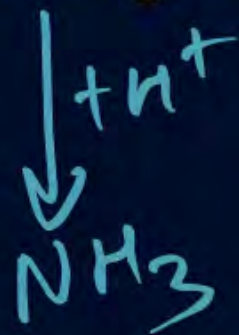
B H_3O^+ and F^- , respectively

C OH^- and F^- , respectively

D H_3O^+ and H_2F^+ , respectively

QUESTION – (AIPMT 2000)

Conjugate acid of NH_2^- is:



QUESTION – (NEET Odisha 2019)

Which of the following cannot act both as Bronsted acid and as Bronsted base?

- A** HSO_4^- $\xrightarrow{-\text{H}^+} \text{SO}_4^{2-}$ $\xrightarrow{+\text{H}^+} \text{H}_2\text{SO}_4$
- B** HCO_3^- $\xrightarrow{-\text{H}^+} \text{CO}_3^{2-}$ $\xrightarrow{+\text{H}^+} \text{H}_2\text{CO}_3$
- C** NH_3 $\xrightarrow{-\text{H}^+} \text{NH}_2^-$ $\xrightarrow{+\text{H}^+} \text{NH}_4^+$
- ☒ **D** HCl $\xrightarrow{-\text{H}^+} \text{Cl}^-$

MEDIC8 Test

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Complete Thermodynamics → Easy

THANK
YOU