YAKEEN NEET 2.0

2026

Redox Reaction

Physical Chemistry | Lecture -04

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Topics to be covered

1. Revision of Last Class
2. n factor calculation
3. Balancing of redox reaction
4. MEDICS Test no 4
5. Magarmach Practice Questions (MPQ) & Home work from modules

Rules to Attend Class

1. Always sit in a peaceful environment with headphone and be ready with your copy and pen.
2. Never ever attend a class from in between or don't join a live class in the middle of the chapter.
3. Make sure to revise the last class before attending the next class & always complete your Magarmach Practice Questions.
4. Never ever engage in chat whether live or recorded on the topic which is not being discussed in current class as by doing so u can be blocked by the admin team or your subscription can be cancelled.

Rules to Attend Class

5. Try to make maximum notes during the class if something is left then u can use the notes pdf after the class to complete the remaining class.
6. Always ask your doubts in doubt section to get answer from faculty. Before asking any doubt please check whether same doubt has been asked by someone or not.


There is one big flaw in your Preparation that's name is Backlog ? What do we say to Backlog ?
NOT TODAY !!!

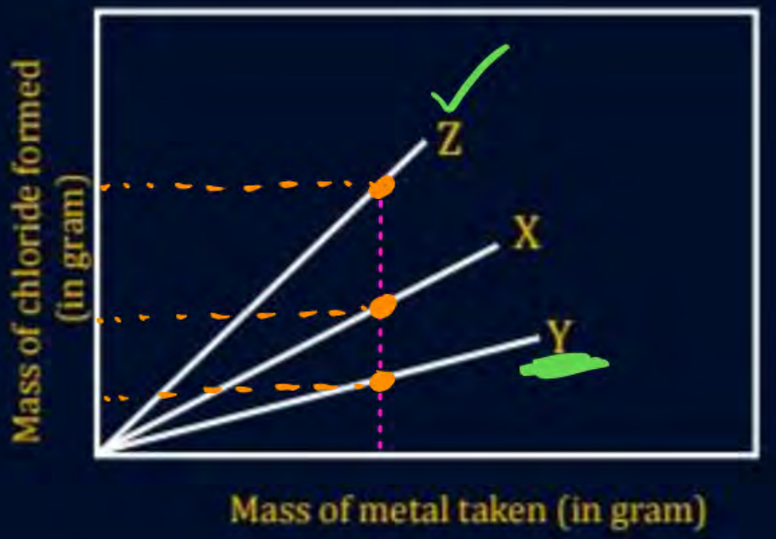
QUESTION

2nd group ↑ metals

$M + Cl_2 \rightarrow MCl_2$.

Alkaline earth metals X, Y, Z on reaction with Cl_2 form chloride. Graph between amount of metal taken (along X-axis) and amount of chloride formed (along Y-axis) is of the type.

MCl_2.



M

Thus, atomic masses of metals X, Y and Z are in order

$Z < X < Y$

(A) $X < Y < Z$

(B) $Y < X < Z$

(C) $Z < X < Y$ ✓

(D) $Z < Y < X$

^{24}Mg ^{40}Ca

$W_{Mg} = W_{Ca} = 240\,g$

$n_{Mg} = \frac{240}{24} = 10$ | $n_{Ca} = \frac{240}{240} = 6$

$10\,Mg + 10\,Cl_2 \rightarrow 10\,MgCl_2$ | $6\,Ca + 6\,Cl_2 \rightarrow 6\,CaCl_2$

$10 \times 95 = 950\,g$ | $6 \times 111 = 666\,g$

MEDICS

Mastery
Checks your grasp over NEET-level concepts

Evaluation
Judging both knowledge and test-smartness

Decision Making
Testing your speed + accuracy under pressure

Intuition
Some answers need gut + logic – can you spot the trick?

Concepts
It's all about strong basics – no shortcuts here

Strategy
The MEDICS test – built for those who heal, hustle, and hope.

Q1. Which of the following statements best defines oxidation?

A. Loss of protons

B. Gain of electrons

C. Loss of electrons

D None of these

Q2. $Mg + Cl_2 \rightarrow MgCl_2$

What is the oxidizing agent?

A. Mg B. Cl_2

C. $MgCl_2$ D. Both Mg and Cl_2

Q3. Which species undergoes reduction in the following reaction?

A. CO B. Fe

C. Fe_2O_3 D. CO_2

$Fe_2O_3 + CO \rightarrow Fe + CO_2$

Q4. What is the oxidation number of sulfur in H_2SO_4?

A. +2 B. +4

C. +6 D. −2

H_2SO_4

$2 + x - 8 = 0$

$x = +6$

Q5. Which of the following is a reducing agent?

A. H_2O B. Cl_2

C. Na D. HCl

Q6. When $KMnO_4$ is reduced in acidic inedium, the change in oxidation number of Mn is: $Mn^{+7}O_4^- \rightarrow Mn^{2+}$

A. +7 to +4 B. +7 to 0

C. +7 to +2 D. +6 to 3

Q7. The n-factor of H_2SO_4 in its reaction with NaOH is:

A. 1 B. 2

C. 0 D. 4

$H_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2H_2O$

Q8. In which of the following does oxidation occur?

A. $Zn \rightarrow Zn^{2+} + 2e^-$

B. $Cu^{2+} + 2e^- \rightarrow Cu$

C. $O_2 + 4e^- \rightarrow 2O^{2-}$

D None of these

Revision of Last class

Synthesis

$A + B \rightarrow AB$

$C + O_2 \rightarrow CO_2$

Decomposition

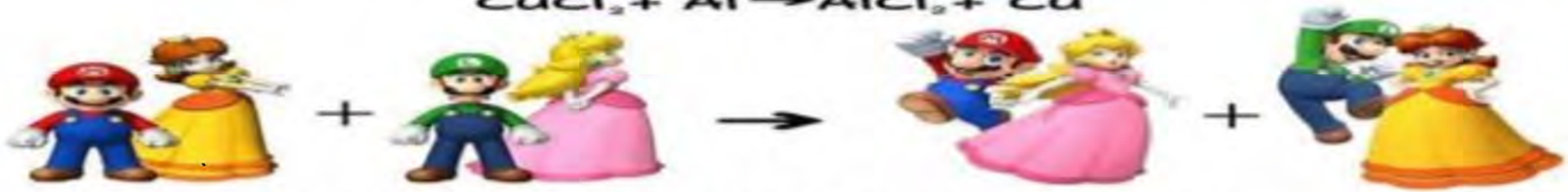
$AB \rightarrow A + B$

$ClNa \rightarrow Cl + Na$

Single Replacement

$A + BC \rightarrow AB + C$

$CuCl_2 + Al \rightarrow AlCl_3 + Cu$



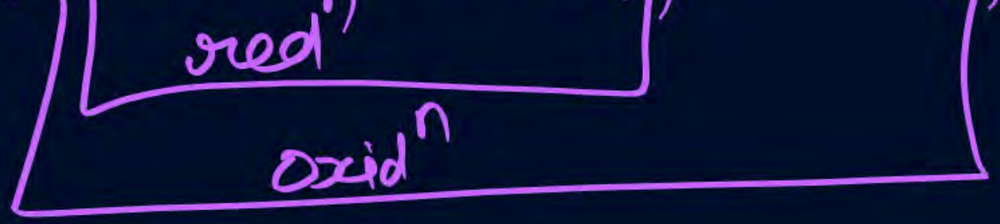
Double Replacement

$AB + CD \rightarrow AD + CB$

$NaCl + AgNO_3 \rightarrow NaNO_3 + AgCl$

Disproportionation.

$$\overset{-1}{2H_2O_2} \longrightarrow \overset{-2}{2H_2O} + \overset{0}{1O_2}$$



Comproportionation

Reverse of disproportionation.

$$\overset{-2}{2H_2O} + \overset{0}{1O_2} \longrightarrow \overset{-1}{2H_2O_2}$$

n factor calculation

① $\underset{+p\ +q}{A_x B_y} \rightarrow \underset{+x\ +q}{A_z B_w}$

$n_f = x|p - x|$

② $2KMnO_4 + 16\overset{-1}{H}\overset{-1}{Cl} \rightarrow 2K\overset{-1}{Cl} + 2Mn\overset{-1}{Cl_2} + 5\overset{0}{Cl_2} + H_2O$

$n_f = \frac{10|-1-0|}{16} = \frac{10}{16} = \frac{5}{8}$

Eq. mass of $HCl = \frac{36.5 \times 8}{5}$

③ $\underset{A_x B_y}{\overset{+p \; +q}{}} \longrightarrow \underset{A_Z B_w}{\overset{+r \; +s}{}}$

$\downarrow$

Total $n_f = \Sigma n_p$

$O_2^+ PtF_6^-$

$\downarrow$

$1 + x - 6 = 0$

$x = +5$

#MET

(H) Case-4

(a) Molecule diff. elements oxidised & reduced.

(b) n_f Total $= \left| \Sigma n_f \text{ oxid}^n - \Sigma n_f \text{ red}^n \right| \neq 0$

if $= 0$ then n_f Total $= \Sigma n_f \text{ oxid}^n$ or $\Sigma n_f \text{ red}^n$

Q. $\overset{+1\ +5\ -2}{KClO_3} \xrightarrow{\Delta} \overset{+1\ -1}{KCl} + \overset{0}{O_2}$

n_f Total $= 6$

red^n

oxid^n

$n_f \text{ oxid}^n = 3|-2-0| = 6$

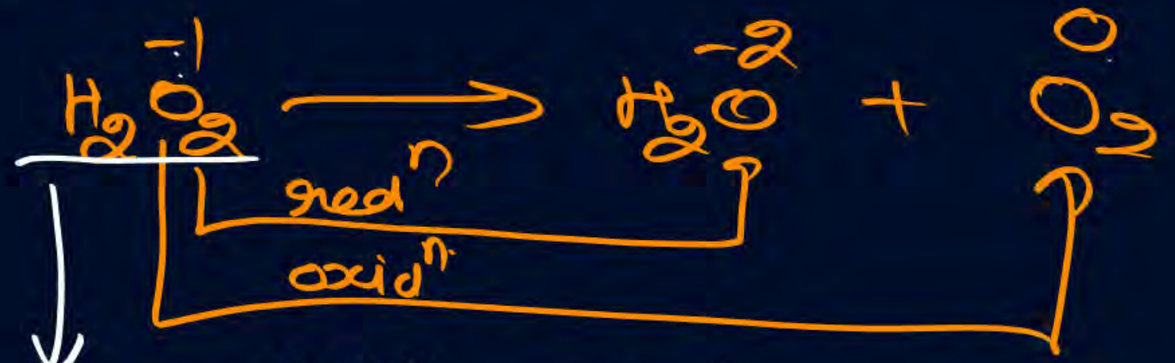
$n_f \text{ red}^n = 1|5-(-1)| = 6$

(5) Disproportionation ^{n}

MIT

$$n_f \text{ Total} = \frac{n_f \text{ oxid}^{n} \times n_f \text{ red}^{n}}{n_f \text{ oxid}^{n} + n_f \text{ red}^{n}}$$

Q =



$$n_f \text{ Total} = \frac{2 \times 2}{2+2} = \frac{4}{4} = 1$$

$$n_f \text{ ox.} = 2|-1-0| = 2$$

$$n_f \text{ red}^{n} = 2|-1-(-2)| = 2$$

Q $Cl_2 + NaOH \xrightarrow[\text{dil \& Conc.}]{} NaCl^{-1} + NaOCl^{+1} + H_2O$

n_f Total $= \frac{2\times2}{2+2} = \frac{4}{4} = 1$

n_f ox. $= 2|0-(+1)| = 2$

n_f red $= 2|0-(-1)| = 2$

Q $Cl_2 + NaOH \xrightarrow[\text{Hot \& Conc.}]{} NaCl^{-1} + NaClO_3^{+5} + H_2O$

n_f Total $= \frac{10\times2}{10+2} = \frac{20}{12} = \frac{5}{3}$

n_f ox. $= 2|0-(+5)| = 10$

n_f red $= 2|0-(-1)| = 2$

Q. $3Cl_2 + 6NaOH \longrightarrow 5NaCl + 1NaClO_3 + 3H_2O$

(6 NaOH: Hot & Conc.)

Find eq. mass of NaOH in this Δ^n?

Ans

$$geq\ Cl_2 = geq\ NaOH$$

$$3 \times \frac{5}{3} = 6 \times n_f$$

$$n_f = \frac{5}{6}$$

$$E_{NaOH} = \frac{40}{5} \times 6 = 48\ g$$

Q. $P_4^{0} + \text{aq. } NaOH \longrightarrow \overset{-3}{P}H_3 + NaH_2\overset{+1}{P}O_2 + H_2O$

$\downarrow$ → eq. mass of $P_4 = \frac{124}{n_f} = \frac{124}{3} = 41.33$

$n_f = \frac{12 \times 4}{12+4} = \frac{48}{16} = 3$

n_f oxid$^n = 4|0-(+1)| = 4$

n_f red$^n = 4|0-(-3)| = 12.$

QUESTION – (AIEEE 2008)

In the following change, $3Fe + 4H_2O \longrightarrow Fe_3O_4 + 4H_2$

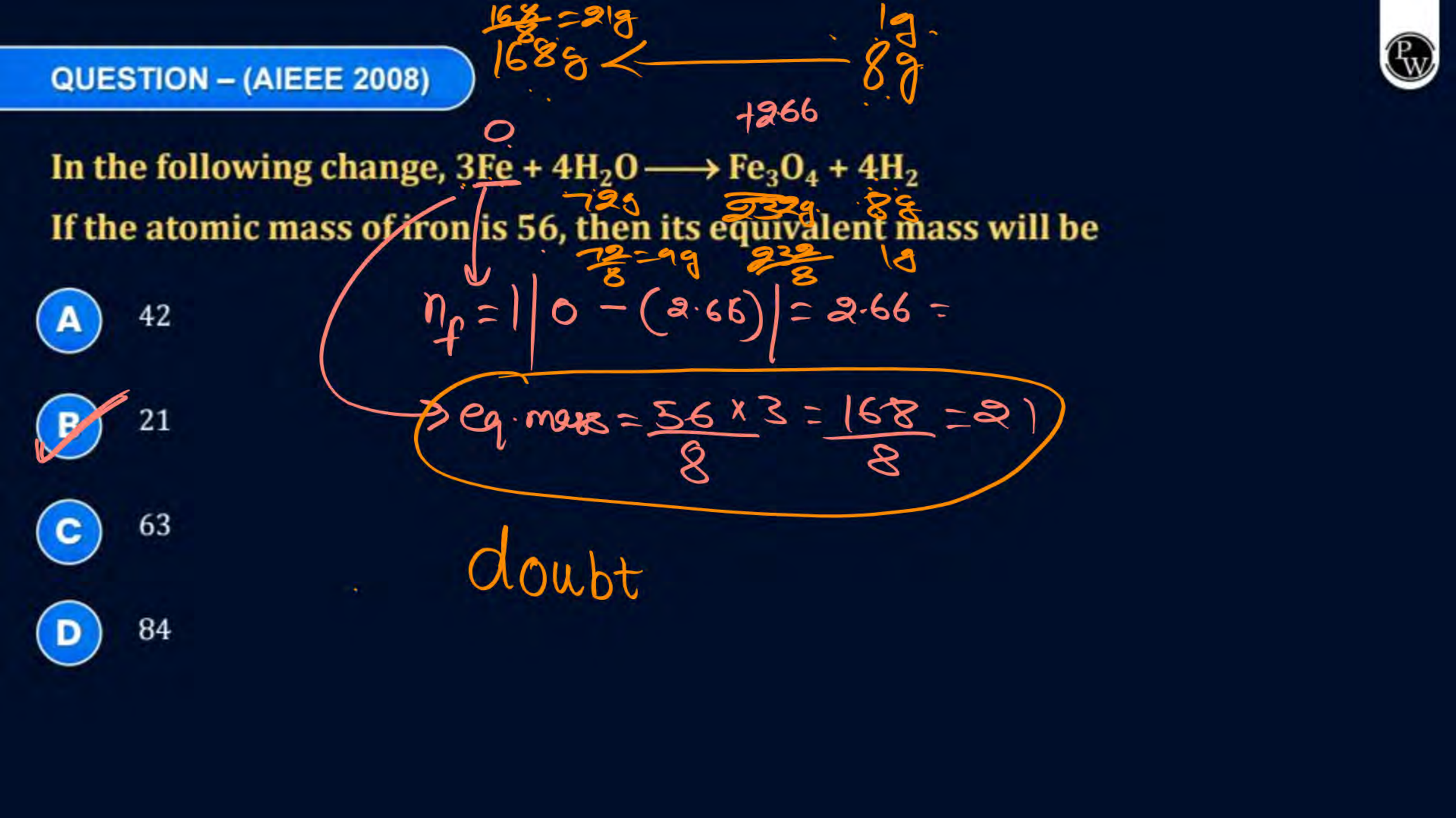
If the atomic mass of iron is 56, then its equivalent mass will be

(A) 42

(B) 21

(C) 63

(D) 84



QUESTION –

The eq. mass of $Na_2S_2O_3$ as reductant in the reaction,

$$\overset{+2}{Na_2S_2O_3} + H_2O + Cl_2 \rightarrow \overset{+6}{Na_2SO_4} + 2HCl + \overset{0}{S}$$

(R.A.)

$n_f = 2|2 - (6)| = 8$

$E = \frac{M}{8}$

(A) M/1

(B) M/2

(C) M/6

(D) M/8

QUESTION –

The eq. mass of iodine in, $\overset{0}{I_2} + 2S_2O_3^{2-} \rightarrow 2I^- + S_4O_6^{2-}$ is

$n_f = 2|0-(-1)| = 2$

$E = \frac{M}{2}$

(A) M

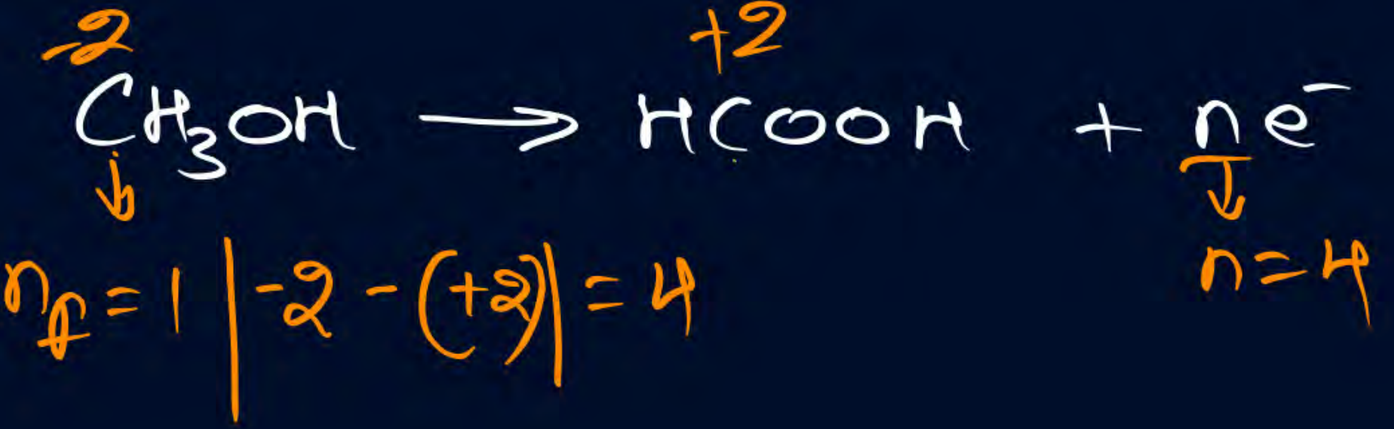
(B) M/2

(C) M/4

(D) None of these

QUESTION –

In the reaction, $CH_3OH \rightarrow HCOOH$ the number of electrons that must be added to the right is



(A) 4

(B) 3

(C) 2

(D) 1

QUESTION –

$n_f = |2-1| = 1$

In the reaction, $2\overset{+2}{CuSO_4} + 4KI \rightarrow \overset{+1}{Cu_2I_2} + 2K_2SO_4 + H_2O + O$

The ratio of equivalent mass of $CuSO_4$ to its molecular mass is

(A) 1/8

(B) 1/4

(C) 1/2

(D) 1 ✓

$$E = \frac{M}{n_f}$$

$$\frac{E}{M} = \frac{1}{n_f}$$

$$\frac{E}{M} = \frac{1}{1}$$

QUESTION –

In alkaline condition $KMnO_4$ reacts as follows,

Basic

$$2\overset{+1}{K}\overset{+7}{Mn}O_4^{-2} + 2KOH \rightarrow 2\overset{+1}{K_2}\overset{+6}{Mn}O_4^{-2} + H_2O$$

Therefore, its equivalent mass will be

$$E = \frac{M}{n_f} = \frac{158}{1}$$

$$n_f = 1|7-6| = 1$$

(A) 31.6

(B) 52.7

(C) 79.0

(D) 158.0 ✓

QUESTION –

In the reaction, $2S_2O_3^{2-} + I_2 \rightarrow S_4O_6^{2-} + 2I^-$;
The eq. mass of $Na_2S_2O_3$ is equal to its

(A) M

(B) M/2

(C) 2 × M

(D) M/6

+2 +2.5

$n_f = 2|2 - 2.5| = 1$

$E = \frac{M}{1}$

$Na_2S_2O_3$

$Na_2S_4O_6$

QUESTION (AIIMS 2018)

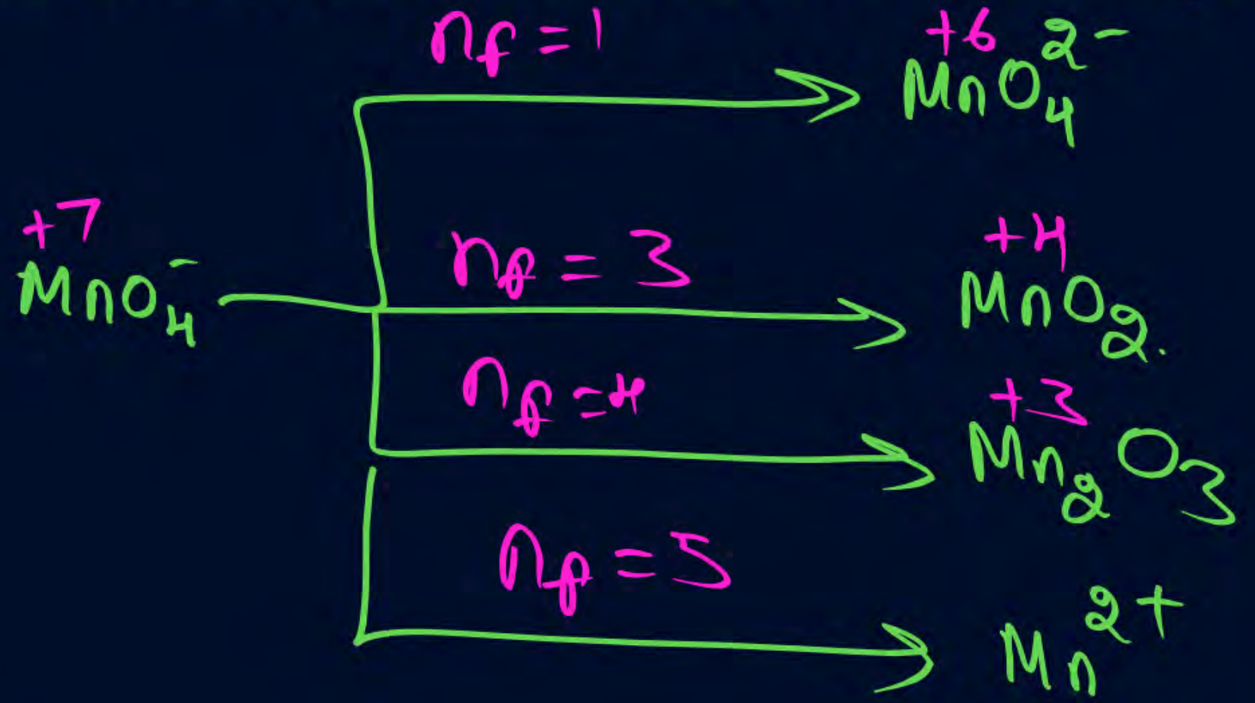
When $KMnO_4$ acts as an oxidizing agent and ultimately forms MnO_4^{2-}, MnO_2, Mn_2O_3 and Mn^{2+}, then the number of electrons transferred in each case respectively is:

(A) 4, 3 1, 5

(B) 1, 5, 3, 7

(C) 1, 3, 4, 5

(D) 3, 5, 7, 1



QUESTION –

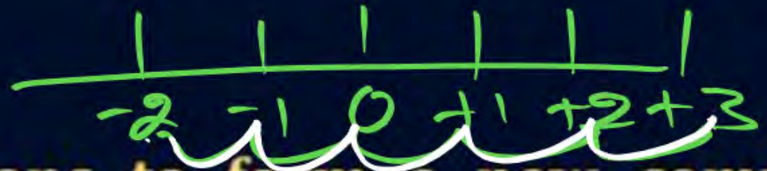
One mole of N_2H_4 loses 10 mole of electrons to form a new compound Y. Assuming that all nitrogen appears in the new compound, what is the oxidation number of nitrogen in Y (there is no change in the oxidation state of hydrogen)?

(A) -3

(B) +3

(C) +5

(D) +1



1 mole N_2H_4 . 10 mole of e^-

N_A molecules — $10N_A$ of e^-

1 — 10 of e^-

2 atoms 10 e^- loss

1 — 5 e^- loss

$\overset{-2}{N_2}H_4 \longrightarrow \underset{(2N)}{Y} + 10e^-$

$2x + 4 = 0$

$2x = -4$

$x = -2$

$n_f = 2|-2 - Z| = 10$

$|-2 - Z| = 5$

$Z = +3$

Balancing of Redox Reactions

#MIT

$\overset{+7\ -2}{MnO_4^-} + \overset{-1}{I^-} + H^+ \rightarrow \overset{+2}{Mn^{2+}} + \overset{0}{I_2} + \overset{+1\ -2}{H_2O}$

① Which is oxidised & reduced.

② $(S.C.)_O = n_{f(R)}$ & $(S.C.)_R = (n_f)_O$

③ Balance all except H & O

④ O_2 balance add H_2O to side deficient

⑤ Balance H

⑥ in neutral & acidic ⇒ add H^+ to side deficient

⑥ in basic medium ⇒ add H_2O to side deficient & equal no. of OH^- to opposite side.

Q=

$$\overset{+7}{MnO_4^-} + 5\overset{-1}{I^-} + 8H^+ \rightarrow \overset{+2}{Mn^{2+}} + \frac{5}{2}\overset{0}{I_2} + 4H_2O$$

R

O

$n_f = |7-2| = 5$

4 ox.

1 Hg.

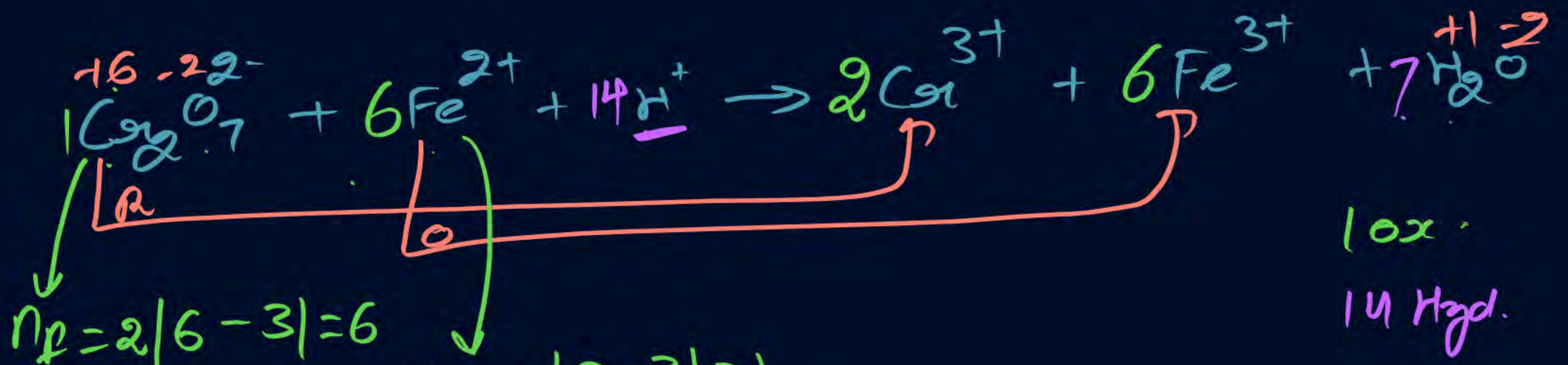
$n_f = |-1-0| = 1$

1 oxc.

8 Hyd.

$$2MnO_4^- + 10I^- + 16H^+ \rightarrow 2Mn^{2+} + 5I_2 + 8H_2O$$

Q11


-16 -22-
Cr_2O_7 + 6 Fe^{2+} + 14 H^+ → 2 Cr^{3+} + 6 Fe^{3+} + 7 H_2O
+1 -2
R
O
$n_f = 2|6-3| = 6$
$n_f = 1|2-3| = 1$
7 ox.
1 Hy
10 ox.
14 Hyd.

Q:

$$2\overset{+6}{Cr_2}\overset{2-}{O_7^{2-}} + 6\overset{+3}{N}O_2^- + 16H^+ \rightarrow 4Cr^{3+} + 6\overset{+5}{N}O_3^- + 8H_2O$$

$n_f = 2|6-3| = 6$

26 ox.

1 Hy

$n_f = 1|3-5| = 2$

19 ox.

16 H

$$Cr_2O_7^{2-} + 3NO_2^- + 8H^+ \rightarrow 2Cr^{3+} + 3NO_3^- + 4H_2O$$

$$\text{g-eq oxid}^n = \text{g-eq red}^n$$

$$(\text{moles})_O (n_f)_O = (\text{moles})_R (n_f)_R$$

$$\frac{(\text{moles})_O}{(\text{moles})_R} = \frac{(n_f)_R}{(n_f)_O}$$

$$\frac{(S.C.)_O}{(S.C.)_R} = \frac{(n_f)_R}{(n_f)_O}$$

QUESTION – (NCERT: PL-246, | JEE Main Feb. 1, 2024 (I))

In acidic medium, $K_2Cr_2O_7$ shows oxidizing action as represented in the half reaction :

$$Cr_2O_7^{2-} + XH^+ + Ye^{\ominus} \rightarrow 2A + ZH_2O$$

X, Y, Z and A are respectively are :

(A) 8, 6, 4 and Cr_2O_3

(B) 14, 7, 6 and Cr^{3+}

(C) 8, 4, 6 and Cr_2O_3

(D) 14, 6, 7 and Cr^{3+}

QUESTION – (NCERT: PL-246 | NV, JEE Main Jan. 31, 2024 (II))

Number of moles of H^+ ions required by 1 mole of MnO_4^- to oxidise oxalate ion to CO_2 is _____.

QUESTION – (NCERT: PL-246 | NV, JEE Main Jan.30, 2024 (I))

$$2MnO_4^- + bI^- + cH_2O \rightarrow xI_2 + yMnO_2 + zOH^-$$

If the above equation is balanced with integral coefficients, the value of z is:

QUESTION – (8th April 1st Shift 2023)

$2IO_3^- + xI^- + 12H^+ \rightarrow 6I_2 + 6H_2O$

What is the value of x?

A 10

B 2

C 12

D 6

QUESTION – (NCERT: PL-246, | JEE Main Jan. 29, 2024 (I))

Chlorine undergoes disproportionation in alkaline medium as shown below :

a $Cl_2(g)$ + b $OH^-(aq)$ → c $ClO^-(aq)$ + d $Cl^-(aq)$ + e $H_2O(l)$

The values of *a, b, c* and *d* in a balanced redox reaction are respectively :

(A) 2, 2, 1 and 3

(B) 1, 2, 1 and 1

(C) 2, 4, 1 and 3

(D) 3, 4, 4 and 2

QUESTION – (31st Aug 2nd Shift 2021)

In which one of the following sets all species show disproportionation reaction?

(A) ClO_4^-, MnO_4^-, ClO_2^- and F_2

(B) MnO_4^-, ClO_2^-, Cl_2 and Mn^{3+}

(C) $Cr_2O_7^{2-}$, MnO_4^-, ClO_2^- and Cl_2

(D) ClO_2, F_2, MnO_4^- and $Cr_2O_7^{2-}$

QUESTION (AIIMS 2017)

Consider the following reaction occurring in basic medium

$2MnO_4^-(aq) + Br^-(aq) \rightarrow 2MnO_2(s) + BrO_3^-(aq)$

How the above reaction can be balanced further?

(A) By adding 2 OH^- ions on right side

(B) By adding one H_2O molecule to left side

(C) By adding $2H^+$ ions on right side

(D) Both (A) and (B)

QUESTION (AIIMS 2005)

In the balanced chemical reaction,

$IO_3^- + aI^- + bH^+ \rightarrow cH_2O + dI_2$

a, b, c and d respectively corresponds to:

A. 5, 6, 3, 3

B. 5, 2, 6, 3

C. 3, 5, 3, 6

D. 5, 6, 5, 5

Home work from modules

Prarambh → Q 44, 45, 48, 49, 50, 51, 53

Prabal → Q 17, 18

Magarmach Practice Questions (MPQ)

QUESTION (NEET 2024)

Which reaction is **NOT** a redox reaction?

(A) $H_2 + Cl_2 \rightarrow 2HCl$

(B) $BaCl_2 + Na_2SO_4 \rightarrow BaSO_4 + 2NaCl$

(C) $Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$

(D) $2KClO_3 + I_2 \rightarrow 2KIO_3 + Cl_2$

QUESTION (NEET 2023)

On balancing the given redox reaction,

$aCr_2O_7^{2-} + bSO_3^{2-}(aq) + cH^+(aq) \rightarrow 2aCr^{3+}(aq) + bSO_4^{2-}(aq) + c/2\ H_2O(l)$

the coefficients a, b and c are found to be respectively:

(A) 8, 1, 3

(B) 1, 3, 8

(C) 3, 8, 1

(D) 1, 8, 3

QUESTION (NEET 2020)

What is the change in oxidation number of carbon in the following reaction?

$CH_4(g) + 4Cl_2(g) \rightarrow CCl_4(l) + 4HCl(g)$

(A) 0 to +4

(B) −4 to +4

(C) 0 to −4

(D) +4 to +4

QUESTION (NEET 2020-Covid)

The oxidation number of the underlined atom in the following species.
Identify the incorrect option.

(A) $\underline{Cl}O_3^-$ is +5

(B) $K_2\underline{Cr}_2O_7$ is +6

(C) $H\underline{Au}Cl_4$ is +3

(D) $Cu_2\underline{O}$ is –1

QUESTION – (AIEEE 2019)

Which of the following reactions are disproportionation reaction?

A. $2Cu^{+} \longrightarrow Cu^{2+} + Cu^{0}$

B. $3MnO_4^{2-} + 4H^{+} \longrightarrow 2MnO_4^{-} + MnO_2 + 2H_2O$

C. $2KMnO_4 \xrightarrow{\Delta} K_2MnO_4 + MnO_2 + O_2$

D. $2MnO_4^{-} + 3Mn^{2+} + 2H_2O \longrightarrow 5MnO_2 + 4H^{+}$

Select the correct option from the following:

A. (A) and (B) only

B. (A), (B) and (C)

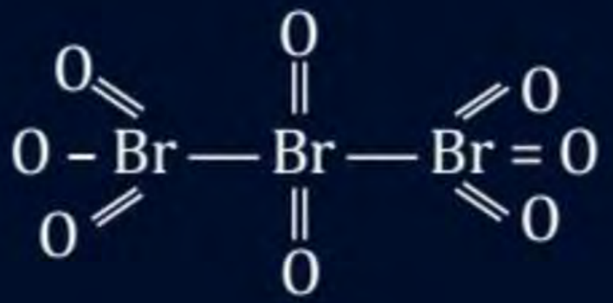
C. (A), (C) and (D)

D. (A) and (D) only

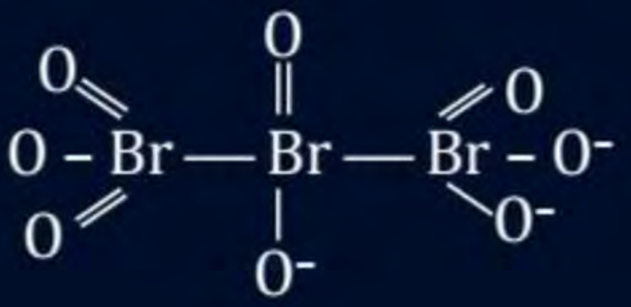
QUESTION – (AIEEE 2019)

The correct statement of tribromooctaoxide is:

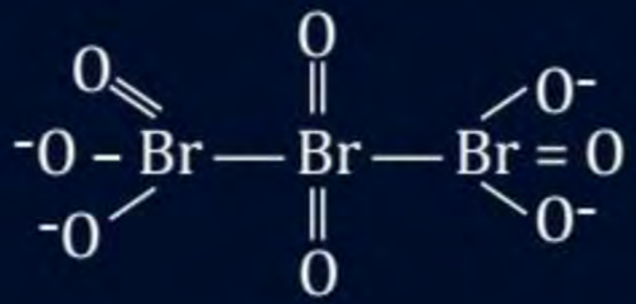
A



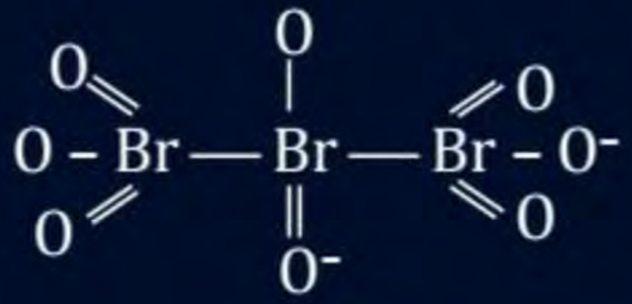
B



C



D



QUESTION (NEET 2018)

For the redox reaction

$MnO_4^- + C_2O_4^{2-} + H^+ \rightarrow Mn^{2+} + CO_2 + H_2O$

The correct coefficients of the reactants for the balanced equation are:

	MnO_4^-	$C_2O_4^{2-}$	H^+
A	16	5	2
B	2	5	16
C	5	16	2
D	2	16	5

QUESTION (NEET 2016 - I)

Hot concentrated Sulphuric acid is a moderately strong oxidizing agent. Which of the following reactions does not show oxidizing behavior?

(A) $C + 2H_2SO_4 \rightarrow CO_2 + 2SO_2 + 2H_2O$

(B) $CaF_2 + H_2SO_4 \rightarrow CaSO_4 + 2HF$

(C) $Cu + 2H_2SO_4 \rightarrow CuSO_4 + SO_2 + 2H_2O$

(D) $3S + 2H_2SO_4 \rightarrow 3SO_2 + 2H_2O$

QUESTION (NEET 2014)

In acidic medium, H_2O_2 changes $Cr_2O_7^{-2}$ to CrO_5 which has two (—O—O—) bonds. Oxidation state of Cr in CrO_5 is:

A +3

B +6

C –10

D +5

QUESTION (NEET 2014)

The oxidation state of Cr in CrO_5 is:

A. −6

B. +12

C. +6

D. +4

QUESTION (AIPMT 2012)

When Cl_2 gas reacts with hot and concentrated sodium hydroxide solution, the oxidation number of chlorine changes from:

A zero to +1 and zero to –5

B zero to –1 and zero to +5

C zero to –1 and zero to +3

D zero to +1 and zero to –3

QUESTION (AIPMT 2009)

Oxidation numbers of P in PO_4^{3-}, of S in SO_4^{2-} and the Cr in $Cr_2O_7^{2-}$ are respectively

(A) +3, +6 and +5

(B) + 5, +3 and +6

(C) –3, +6 and +6

(D) +5, +6 and +6

QUESTION (AIPMT 2003)

The oxidation states of Sulphur in the anions SO_3^{2-}, $S_2O_4^{2-}$ and $S_2O_6^{2-}$ follow the order

(A) $S_2O_6^{2-} < S_2O_4^{2-} < SO_3^{2-}$

(B) $S_2O_4^{2-} < SO_3^{2-} < S_2O_6^{2-}$

(C) $SO_3^{2-} < S_2O_4^{2-} < S_2O_6^{2-}$

(D) $S_2O_4^{2-} < S_2O_6^{2-} < SO_3^{2-}$

QUESTION (AIPMT 2000)

A compound contains atoms of three elements A, B and C. If the oxidation number of A is +2, B is +5, and that of C is –2. The possible formula of the compound is:

(A) $A_2(BC_3)_2$

(B) $A_3(BC_4)_2$

(C) $A_3(B_4C)_2$

(D) ABC_2

QUESTION (AIPMT 1999)

The oxidation number of phosphorus in pyro-phosphoric acid is:

A. +3

B. +1

C. +4

D. +5

QUESTION (1988, 1995)

The oxidation number of chromium in potassium dichromate is:

(A) +6

(B) –5

(C) –2

(D) +2

QUESTION (1994)

Phosphorus has the oxidation state of +3 in

A. Phosphorous acid

B. Orthophosphoric acid

C. Hypophosphorous acid

D. Metaphosphoric acid

QUESTION (1999)

The following redox reaction is balanced by which set of coefficients?

$aZn + bNO_3^- + cH^+ \rightarrow dNH_4^+ + eH_2O + fZn^{2+}$

	a	b	c	d	e	f
A	1	1	10	1	3	1
B	2	2	10	2	3	2
C	4	2	10	1	3	4
D	4	1	10	1	3	4

QUESTION (1994)

In which of the following reactions, there is no change in valency?

A $4KClO_3 \longrightarrow 3KClO_4 + KCl$

B $SO_2 + 2H_2S \longrightarrow 2H_2O + 3S$

C $BaO_2 + H_2SO_4 \longrightarrow BaSO_4 + H_2O_2$

D $3BaO + O_2 \longrightarrow 2BaO_2$

QUESTION (1994)

Which substance serves as a reducing agent in the following reaction?

$14H^+ + Cr_2O_7^{2-} + 3Ni \rightarrow 2Cr^{3+} + 7H_2O + 3Ni^{2+}$

A H_2O

B Ni

C H^+

D $Cr_2O_7^{2-}$

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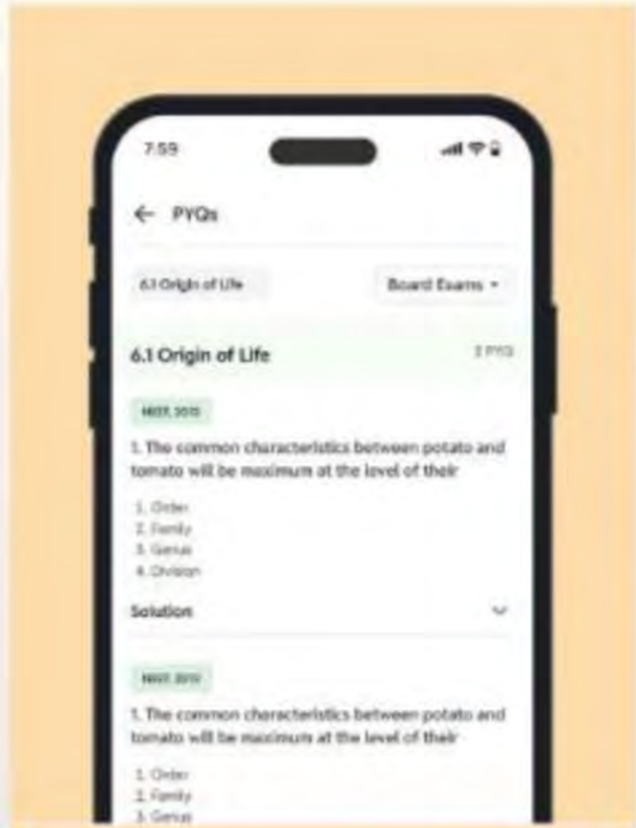
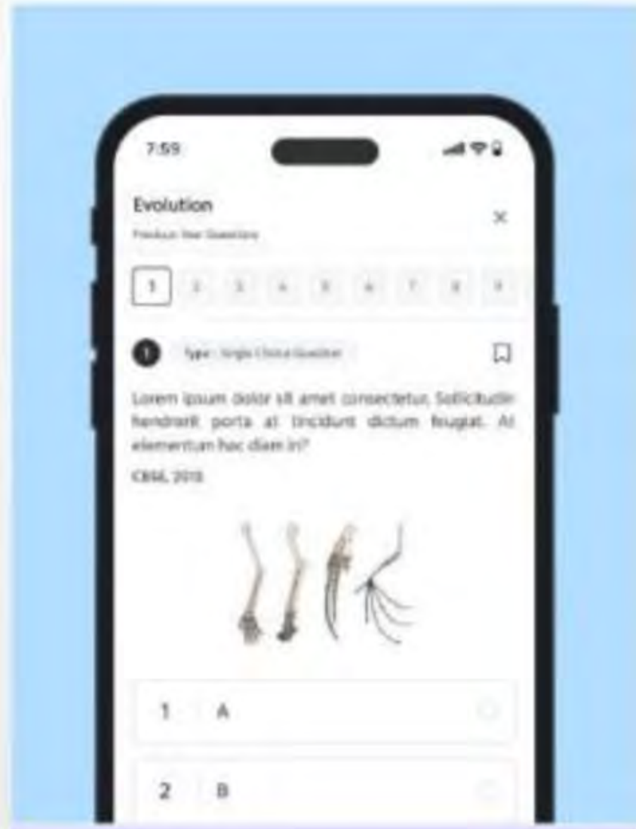


Topic-wise

Previous Year Questions



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Practice Questions to Ace Exams

MEDICS → Test no :- 5

Some Basic Concepts of Chem. → eq. mass
g. eq.
N, law of equivalence.
↓
Moderate.

THANK YOU