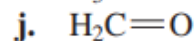
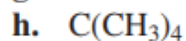
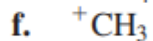
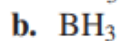
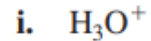
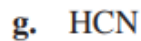
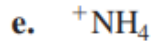
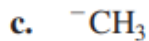
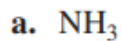
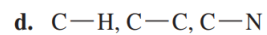
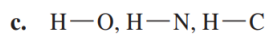
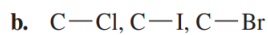
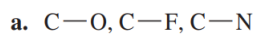


1. What is the hybridization of all the atoms (other than hydrogen) in each of the following?

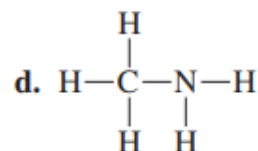
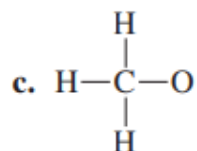
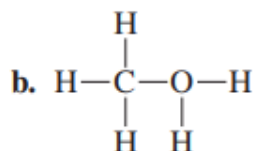
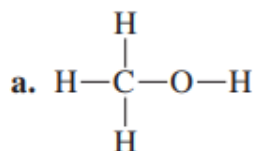


a. \_\_\_\_\_ c. \_\_\_\_\_ e. \_\_\_\_\_ g. \_\_\_\_\_ i. \_\_\_\_\_  
b. \_\_\_\_\_ d. \_\_\_\_\_ f. \_\_\_\_\_ h. \_\_\_\_\_ j. \_\_\_\_\_

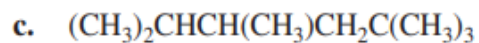
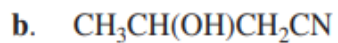
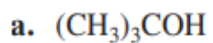
2. Rank the bonds from most polar to least polar.



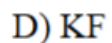
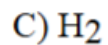
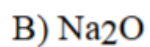
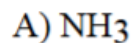
3. Draw the missing lone-pair electrons and assign the missing formal charges for the following:



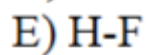
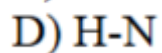
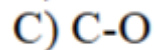
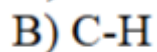
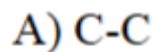
4. Draw a Lewis structure for each of the following:



5. Which of the following contain(s) polar covalent bonds?



6. Which of the following covalent bonds has the largest dipole moment?



7. Which of the following molecules does not exhibit a net dipole moment of zero?

