



$$I_{D3} \approx I_{D4}$$

$$I_{D4} = I_{D3} K$$

$$\frac{I_{D3}}{I_{D4}} = \frac{1}{K}$$

$$V_{GS3} = V_{GS4} + I_{D4} R_S$$

assuming $V_{th3} = V_{th4}$

$$\sqrt{\frac{I_{D3}^2}{\mu_n C_{ox} \frac{W}{L}_3}} + V_{th3} = \sqrt{\frac{2 I_{D4}}{\mu_n C_{ox} \frac{W}{L}_4}} \times V_{th4} (1 + R_S) + V_{th4}$$

$$\sqrt{\frac{I_{D3}^2}{\mu_n C_{ox} \frac{W}{L}_3}}$$

$$= (1 + R_S)$$

$$\sqrt{\frac{2 I_{D4}}{\mu_n C_{ox} \frac{W}{L}_4}}$$

$$\frac{\frac{I_{D3}}{W/L_3}}{\frac{I_{D4}}{W/L_4}} = \sqrt{(1 + R_S)} = \frac{1}{K} = \frac{W/L_4}{W/L_3} \sqrt{(1 + R_S)}$$

$$K = \frac{W/L_3}{W/L_4} \frac{1}{\sqrt{(1 + R_S)}}$$

$$I_{D4} = I_{D3} \times K$$

V_{DD} V_{REF}

$$V_{SD2} = V_{DD} - V_{DS4} - I_{D4} R_S$$

$$I_{D4} = \frac{W}{L} \frac{1}{2} C_{ox} \mu_p (V_{SD2} - V_{th})^2 \propto 2$$

the Relation But with an Increase in

$$\frac{I_{D3}}{I_{D4}} = k_i \rightarrow \text{this reaction is controlled}$$

But I_{D3} controls V_{ref}

If we increase $\frac{W}{L}$ of $\frac{W}{L}_3 \rightarrow V_{ref}$ will Decrease
and k will increase. Because and I_{D4} will increase