

Rate Expression

intro

- reaction mechanism: sequence that reaction goes through
 - most reactions involve multiple steps
 - each step involves collision
 - one step will be slower than others
- Rate expression - constant multiplied by reactants
 - order of reaction refers to exponents of concentrations
 - exponents can only be found experimentally
- $\text{rate} = k[\text{reactants}]^{\text{order}}$