

- Q.1** Soybean seeds are extracted with hexane in a batch extractor. The seeds contain 18.6% oil, 69% solids and 12.4% moisture. At the end of extraction process, deoiled cake is separated from the hexane oil mixture. Deoiled cake analysis yields 0.8% oil, 87.7% solids and 11.5% moisture, find the % recovery of the oil.
- Q.2** The average MW of a flue gas sample is calculated by two different engineers. One engineer uses the correct MW for N_2 and determines the average MW to be 30.08. The other engineer, using an incorrect value for the MW of N_2 (i.e 14), calculate the average MW to be 18.74.
- (a) Calculate the volume% of N_2 in the flue gas
- (b) If the remaining components of the flue gases are CO_2 and O_2 , calculate the volume% of each.
- Q.3** A 100 kg mixture of 27.8% of acetone (A) and 72.2% of chloroform (B) by weight is to be extracted with a mixed solvent. The solvent contains unknown composition of water (S1) and acetic acid (S2). The mixture, of the original mixture and solvent, is shaken well, allowed to attain equilibrium and then separated into 2 layers. The composition of the 2 layers is:-

	A	B	s1	s2
Upper layer (mass%)	7.5	3.5	57.4	31.6
Lower layer (mass%)	20.3	67.3	2.8	9.6

Calculate (a) the quantities of the 2 layers (b) the weight ratio of the mixed solvent to the original mixture (c) the composition of the mixed solvent.

- Q.1** Soybean seeds are extracted with hexane in a batch extractor. The seeds contain 18.6% oil, 69% solids and 12.4% moisture. At the end of extraction process, deoiled cake is separated from the hexane oil mixture. Deoiled cake analysis yields 0.8% oil, 87.7% solids and 11.5% moisture, find the % recovery of the oil.

Solution Tutorial Sheet 4

Q.1 Basis = 100 kg of seeds

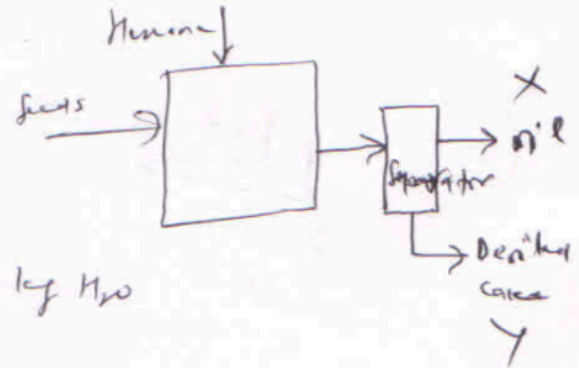
18.6 kg oil; 69 kg solids, 12.4 kg H₂O

The material balance (solids).

$$69 \text{ kg} = Y \times 0.877 \Rightarrow Y = \underline{\underline{78.68 \text{ kg}}}$$

$$\begin{aligned} \text{Oil in oil from} &= \text{Oil in feed} - \text{Oil in cake} \\ &= 18.6 - 0.63 = 17.97 \text{ kg} \end{aligned}$$

$$\% \text{ Recovery} = \frac{17.97}{18.6} = 96.6\%$$



Q.2 The average MW of a flue gas sample is calculated by 2 different engineers. One engineer uses the correct MW for N_2 and determines the average MW to be 30.08. The other engineer, using an incorrect value for the MW of N_2 (i.e 14), calculate the average MW to be 18.74.

(a) Calculate the volume% of N_2 in the flue gas

(b) If the remaining components of the flue gases are CO_2 and O_2 , calculate the volume% of each.

Q.2. c) Let the mixture contain n number of gases one of which is N_2 .

$$\text{Avg. mol. wt. of mix} = x_1 MW_1 + \dots + x_{N_2} MW_{N_2}$$

$$30.08 = x_1 MW_1 + \dots + x_{N_2} (28) \quad \text{--- (I)}$$

$$18.74 = x_1 MW_1 + \dots + x_{N_2} (14) \quad \text{--- (II)}$$

from (I) & (II) $x_{N_2} = 0.81$

~~(b) $30.08 = x_{CO_2} + \dots$~~

$$\begin{aligned} \text{(b)} \quad 30.08 &= 0.81 \times 28 + x_{CO_2} \times 44 + x_{O_2} \times 32 \\ &= 0.81 \times 28 + x_{CO_2} \times 44 + (1 - 0.81 - x_{CO_2}) \times 32 \end{aligned}$$

$$0.11 = x_{CO_2} \Rightarrow x_{CO_2} = 11\%$$

$$\therefore x_{O_2} = 1 - 0.81 - 0.11 = 0.08 \Rightarrow 8\%$$

Q.3 A 100 kg mixture of 27.8% of acetone (A) and 72.2% of chloroform (B) by weight is to be extracted with a mixed solvent. The solvent contains unknown composition of water (S1) and acetic acid (S2). The mixture, of the original mixture and solvent, is shaken well, allowed to attain equilibrium and then separated into 2 layers. The composition of the 2 layers is:-

	A	B	s1	s2
Upper layer	7.5	3.5	57.4	31.6
Lower layer	20.3	67.3	2.8	9.6

Calculate (a) the quantity of the 2 layer (b) the weight ratio of the mixed solvent to the original (c) the composition of the mixed solvent.

③

② M.B. on A: $27.8 = UL \times 0.075 + LL \times 0.203$
M.B. on B: $72.2 = UL \times 0.035 + LL \times 0.673$
 $UL = 93.5 \text{ kg}, LL = 102.4 \text{ kg}$

⑥ $MS = UL + LL - \text{Feed} = 93.5 + 102.4 - 100 = 95.9 \text{ kg}$
Mass ratio = $\frac{95.9}{100} = 0.959$

⑦ $x_{s1} \times 95.9 = UL \times 0.574 + LL \times 0.024$
 $x_{s1} = \frac{57.67 + 2.87}{95.9} = 0.611$
 $x_{s2} = 1 - 0.589 = 0.411$