

# Carnot Cycle Exercise

**Vincent Edwards**

Mt. San Antonio College, Physics 4B, CRN 42240  
March 27, 2023

$$T_H = 490 \text{ K}$$

$$V_c = 1.90 \times 10^{-3} \text{ m}^3$$

## 1. Purpose

The goal of the exercise is to perform various calculations related to the Carnot cycle.

## 2. Given

- $T_H = 490 \text{ K}$
- $T_C = 300 \text{ K}$
- $P_c = 1.01 \times 10^5 \text{ Pa}$
- $V_c = 1.90 \times 10^{-3} \text{ m}^3$
- $Q_{a \rightarrow b} = 300 \text{ J}$
- $\gamma = 1.40$
- d.o.f. = 5
- $C_v = \frac{5}{2}R$
- $C_p = \frac{7}{2}R$

## 3. Derivations

### 3.1. Temperature–Volume Relationship for Adiabatic Process

$$\begin{aligned}
 P_i V_i^\gamma &= P_f V_f^\gamma \\
 P_i V_i V_i^{\gamma-1} &= P_f V_f V_f^{\gamma-1} \\
 nRT_i V_i^{\gamma-1} &= nRT_f V_f^{\gamma-1} \\
 T_i V_i^{\gamma-1} &= T_f V_f^{\gamma-1}
 \end{aligned}$$

### 3.2. Work by Gas for Isothermal Process

$$\begin{aligned}
 W &= \int_{V=V_i}^{V=V_f} P dV \\
 W &= \int_{V_i}^{V_f} \frac{nRT}{V} dV \\
 W &= nRT \ln(V) \Big|_{V_i}^{V_f} \\
 W &= nRT (\ln(V_f) - \ln(V_i)) \\
 W &= nRT \ln \left( \frac{V_f}{V_i} \right)
 \end{aligned}$$

## 4. Results

### 4.1. Moles of Gas ( $n$ )

$$\begin{aligned}
 P_c V_c &= n R T_c \\
 n &= \frac{P_c V_c}{R T_c} \\
 n &= \frac{P_c V_c}{R T_C} \\
 n &= 0.0770 \text{ mol}
 \end{aligned}$$

### 4.2. Pressure ( $P_b$ ) and Volume ( $V_b$ ) at $b$

$$\begin{aligned}
 T_b V_b^{\gamma-1} &= T_c V_c^{\gamma-1} \\
 V_b &= V_c \left( \frac{T_c}{T_b} \right)^{\frac{1}{\gamma-1}} \\
 V_b &= V_c \left( \frac{T_C}{T_H} \right)^{\frac{1}{\gamma-1}} \\
 V_b &= 5.57 \times 10^{-4} \text{ m}^3
 \end{aligned}$$

$$\begin{aligned}
 P_b V_b &= n R T_b \\
 P_b &= \frac{n R T_b}{V_b} \\
 P_b &= \frac{n R T_H}{V_b} \\
 P_b &= 5.62 \times 10^5 \text{ Pa}
 \end{aligned}$$

### 4.3. Pressure ( $P_a$ ) and Volume ( $V_a$ ) at $a$

$$\begin{aligned}
 \Delta U_{a \rightarrow b} &= Q_{a \rightarrow b} - W_{a \rightarrow b} \\
 0 &= Q_{a \rightarrow b} - n R T_H \ln \left( \frac{V_b}{V_a} \right) \\
 \ln \left( \frac{V_b}{V_a} \right) &= \frac{Q_{a \rightarrow b}}{n R T_H} \\
 \frac{V_b}{V_a} &= e^{Q_{a \rightarrow b} / (n R T_H)} \\
 V_a &= V_b e^{-Q_{a \rightarrow b} / (n R T_H)} \\
 V_a &= 2.14 \times 10^{-4} \text{ m}^3
 \end{aligned}$$

$$P_a V_a = n R T_a$$

$$P_a = \frac{nRT_a}{V_a}$$

$$P_a = \frac{nRT_H}{V_a}$$

$$P_a = 1.46 \times 10^6 \text{ Pa}$$

**Table 1.** Pressure, Volume, and Temperature for Key Points  
 Note:  $T_a = T_b = T_H$  and  $T_c = T_d = T_C$

Point	$P$ (Pa)	$V$ (m <sup>3</sup> )	$T$ (K)
a			490
b	$5.62 \times 10^5$	$5.57 \times 10^{-4}$	490
c	$1.01 \times 10^5$	$1.90 \times 10^{-3}$	300
d			300

## 5. Conclusion

## 6. Citations

- [1] Karen Schnurbusch, *Physics 4B Lab Book*, Mt. San Antonio College, 2023, pp. 35-38.
- [2] Karen Schnurbusch, *Physics 4B Equations*, Mt. San Antonio College, 2023, pp. 1-3.