

Specific Heat Capacity of Metals

PHYS 442

Leonid Kozlov

November 17, 2015

Date Performed: November 13, 2015
Partners: Whole class
Instructor: Dr. Shultz

1 Objective

The objective of this experiment is to measure the specific heat capacity of three different samples of unknown metal and to compare those with the accepted values. The samples consist of aluminum, zinc and copper.

2 Definitions

Heat Heat is the measure of the internal kinetic energy of a substance.

Temperature Temperature is a measure of the kinetic energy of a particle. It is the degree or intensity of heat in a substance. Celcius is a unit of temperature. One degree Celcius represents the temperature change of one gram of water when 2.39×10^{-5} Joules of heat is added to it.

Specific Heat Capacity The specific heat capacity is the energy transferred to one kilogram of substance causing its temperature to increase by one degree Celcius. Homer and Bowen-Jones (2014)

Thermal Equilibrium Thermal equilibrium is a condition where two substances in physical contact with each other exchange no net heat energy. Substances in thermal equilibrium are at the same temperature.

3 Theory

The change in the internal energy of an object or substance is equal to the product of the mass and the specific heat capacity and the change in temperature.

$$\Delta U = mC_p\Delta T$$

When water and the metal samples are in thermal equilibrium the change in heat of the water is equal in magnitude to the change in heat of the metal.

$$\Delta U_{metal} = \Delta U_{water}$$

From this relationship we may derive a formula for the specific heat capacity of the metal sample given the mass of metal, mass of water, change in temperature of the water, change in temperature of the metal and the specific heat capacity of water.

$$m_{metal}C_{metal}\Delta T_{metal} = m_{water}C_{water}\Delta T_{water}$$

$$C_{metal} = \frac{m_{water}}{m_{metal}} \frac{\Delta T_{water}}{\Delta T_{metal}} C_{water}$$

4 Materials

- Kettle
- Aluminum, zinc and copper samples
- styrofoam cups
- graduated cylinder
- scale
- thermometer
- tongs
- flask of water

5 Method

- a. Weigh the samples and record
- b. Measure 350 ml of water in graduated cylinder and transfer to styrofoam cup
- c. Measure the initial temperature of the water
- d. Boil water and add metal samples to kettle
- e. Use tongs to transfer a sample to the cup with water
- f. Place thermometer in cup, cover it, stir and record equilibrium temperature
- g. Repeat steps b-f for each sample

6 Data

Metal	Mass Metal	Mass Water	Temp Water Initial	Temp Final
Aluminum	90.6 g	22.5 Celcius	26.3 Celcius	
Zinc	64.1 g	22.9 Celcius	24.4 Celcius	
Copper	203.0 g	300 g	22.5 Celcius	26.2 Celcius

Table 1: Experimental data

Material	Specific Heat Capacity
Water	4180 J/kg. $^{\circ}$ C
Aluminum	900 J/kg. $^{\circ}$ C
Zinc	380 J/kg. $^{\circ}$ C
Copper	387 J/kg. $^{\circ}$ C
Iron	452 J/kg. $^{\circ}$ C
Steel	452 J/kg. $^{\circ}$ C
Lead	128 J/kg. $^{\circ}$ C
Silver	230 J/kg. $^{\circ}$ C

Table 2: Known specific heat capacities

7 Example Calculations

This is the calculation for the specific heat capacity of copper.

$$C_{metal} = \frac{m_{water}}{m_{metal}} \frac{\Delta T_{water}}{\Delta T_{metal}} C_{water}$$
$$\Delta T_{water} = 24.8 - 20.8 = 4.0 \text{ Celcius}$$
$$\Delta T_{metal} = 100 - 24.8 = 75.2 \text{ Celcius}$$
$$C_{metal} = \frac{0.350 \text{ kg}}{0.203 \text{ kg}} \frac{4.0 \text{ Celcius}}{75.2 \text{ Celcius}} 4180 \text{ J/kg} \cdot ^\circ \text{C} = 383.3 \text{ J/kg} \cdot ^\circ \text{C}$$

The percent error is calculated as follows.

$$Error = \frac{387 - 383.3}{387} = 0.944\%$$

8 Results

Material	Measured C_p	Percent Error
Aluminum	832 J/kg \cdot $^\circ$ C	7.5
Zinc	453 J/kg \cdot $^\circ$ C	18%
Copper	387 J/kg \cdot $^\circ$ C	6.7%

Table 3: Calculated specific heat capacities

9 Discussion of Error

Errors occur because of heat loss during the experiment. Some heat is lost while the object was transported from kettle to water, which change in temperature was measured. The other source of heat loss was the air and cup, which absorbed some heat and decreased final temperature of an object and water, which temperature was measured.

10 Conclusion

The experiment helped to successfully determine specific heat capacities of the objects and determine objects' material, based on their heat capacities. There was some inaccuracy, however, it was not significant and did not affect results a lot, so the experiment was useful. Inaccuracy was minimized by the use of cups to block the access of air to the water. Cups had low heat capacity and did not absorb a lot of heat. At the end the goal was achieved and all objects' material was determined.

References

Homer, D. and Bowen-Jones, M. (2014). *Physics*. Oxford, 1st edition.