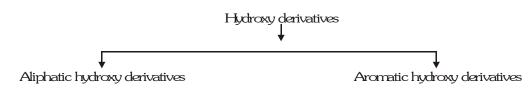
# ALCOHOL, PHENOL & ETHER

# ALCOHOL



(I) Aliphatic hydroxy derivatives :

Hydroxy derivatives in which -OH is directly attached to sp<sup>3</sup> C (Alcoholic compounds).

(II) Aromatic hydroxy derivatives :

Hydroxy derivatives in which -OH is directly attached to  $sp^2 C$  or benzene ring (Phenolic compounds).

Aliphatic hydroxy derivatives :

(a) Classification according to number of -OH groups :

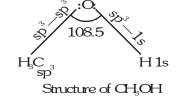
- Monohydric [one -OH]  $\longrightarrow$  CH<sub>3</sub>CH<sub>9</sub>-OH (i)
- $[two -OH] \longrightarrow CH_2-CH_2$ Dihydric (ii)
- Trihydric (iiii)
- Polyhydric [n -OH] (iv)

(b) Classification according to nature of carbon:

- s or 2 alcohol  $\longrightarrow$   $(CH_3)_2CH OH$ t or 3 alcohol  $\longrightarrow$   $(CH_3)_3C OH$

Structure of alcohol:

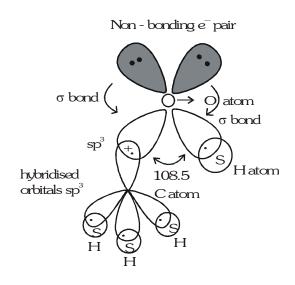
> Alcohols are bent molecules. The carbon atom (linked with 'O' atom of -OH group) is sp3 hybridised. The central 'O' atom is also in sp<sup>3</sup> state of hybridisation. The bond angle is  $108.5^0$  . In  $\,\mathrm{sp^3}$  hybridisation of O -  $2\mathrm{s^2,2p_x^{\phantom{x}2}}\,\,2\mathrm{p_v^{\phantom{y}1}}\,\,2\mathrm{p_z^{\phantom{z}1}}$ orbitals hybridised to form sp3 orbitals



$$\begin{array}{ccc} O & \rightarrow & \uparrow \downarrow & \uparrow \downarrow & \uparrow & \uparrow \\ & 2s & 2p_{x} & 2p_{y} & 2p_{z} \\ & sp^{3} \text{ hybridisation} \end{array}$$

In these four orbitals two containing one electron each and two containing two electrons each. Orbitals containing two electrons do not take part in bonding. Other two half filled orbitals form  $\sigma$ bond with s-orbitals of H -atom and hybridised orbital of C-atom (O-C).

Due to lone pair effect the bond angle of tetrahedral oxygen atom is lesser than normal tetrahedral structure (109028').



## MONOHYDRIC ALCOHOL

#### ☐ General methods of preparation :

(a) From alkanes (By oxidation) :

$$(CH_3)_3$$
  $C-H$   $\xrightarrow{H^{\oplus}/KMnO_4}$   $(CH_3)_3$   $C-OH$ 

(b) From alkenes:

(i) By hydration:

$$CH_3-CH=CH_2 \xrightarrow{H^{\oplus}} CH_3-CH-CH_3$$

$$OH$$

$$OH$$

(ii) By hydroboration oxidation :

$$CH_3-CH=CH_2 \xrightarrow{BH_3 \atop H_2O_2/HO^{\Theta}} CH_3-CH_2-CH_2 (1 \text{ alcohol})$$

(iii) By oxymercuration demercuration :

$$CH_{3}-CH=CH_{2}\xrightarrow{\text{(i)}Hg(OAc)_{2},H_{2}O} CH_{3}-CH-CH_{3}$$

$$OH$$

$$OH$$

(c) From alkyl halides (By hydrolysis):

$$CH_3-CH_2-Cl$$
  $\xrightarrow{\text{Aq. KOH}}$   $CH_3CH_2-OH$ 

(d) From carbonyl compounds (By reduction):

$$>C = O$$
 Reducing agent  $>CH-OH$ 

• Reducing agents may be,

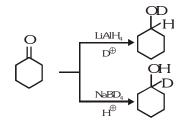
LiAlH<sub>4</sub>/H<sup>⊕</sup>, NaBH<sub>4</sub>/H<sup>⊕</sup>

Na + EtOH [Bouveault-blanc Reduction]

NaH [Darzen reduction]

 $Ni/H_2$ 

♦ Mechanism :



$$\text{CH}_{3}\text{-CH}\!=\!\text{CH}\text{-CHO} \xrightarrow{\text{LiAlH}_{4}} \text{CH}_{3}\text{-CH}\!=\!\text{CH}\text{-CH}_{2}\text{-OH}$$

Crotonaldehyde

$$Ph-CH = CH-CHO \xrightarrow{\text{LiAlH}_4} Ph-CH_2-CH_2-CH_2-OH_2$$

Cinnamaldehyde

(e) From ethers:

(f) From acid and derivatives (By reduction) :

- (g) From esters (By hydrolysis):
  - (i) By alkaline hydrolysis:

$$\begin{array}{ccc} R-C-OR & \xrightarrow{NaOH} & R-C-ONa+R-OH \\ \parallel & \parallel & & \parallel \\ O & & O \end{array}$$

(ii) By acidic hydrolysis:

This reaction is reversible reaction and it's order is 1 and it is also called Pseudo-Unimolecular reaction.

(h) From p-amines:

$$R-NH_{2} \xrightarrow{NaNO_{2}+HCl} R-OH + N_{2} + H_{2}O$$

$$CH_{3}CH_{2}-NH_{2} \xrightarrow{HNO_{2}} CH_{3}CH_{2}-OH + N_{2} + H_{2}O$$

♦ Mechanism :

$$CH_{3}CH_{2} - NH_{2} \xrightarrow{NaNO_{2}+HCl} CH_{3}CH_{2} - \stackrel{\oplus}{N_{2}}\stackrel{\Theta}{Cl} \longrightarrow CH_{3}\stackrel{\oplus}{CH_{2}} + N_{2} + \stackrel{\Theta}{Cl}$$

$$(Unstable)$$

$$\begin{array}{c} \overset{\circ}{\text{OH}} & \text{CH}_{3}\text{CH}_{2} \text{--OH} \text{ [major]} \\ \overset{\circ}{\text{O}} & \text{CH}_{3}\text{CH}_{2} \text{--C} \\ \overset{\circ}{\text{O}} & \text{CH}_{3}\text{CH}_{2} \text{--O-N=O} \\ & \overset{\circ}{\text{O}} & \text{CH}_{3}\text{CH}_{2} \text{--O-N=O} \\ & \overset{\circ}{\text{O}} & \text{CH}_{3}\text{CH}_{2} \text{--O-N=O} \\ & \overset{\circ}{\text{O}} & \text{CH}_{3}\text{CH}_{2} \text{--O-N=O} \end{array}$$

Inter mediate is carbocation so rearrangement may be possible.

Ex. 
$$CH_3CH_2CH_2-NH_2 \xrightarrow{NaNO_2+HCl}$$
 ?

#### Sol. Mechanism:

- (i) From Grignard reagent :
  - (i) p-alcohol:

(ii) s-alcohol:

#### (iii) t-alcohol:

#### Physical properties:

- $\mathrm{C}_1$  to  $\mathrm{C}_{11}$  are colourless liquids and high alcohols are solids. (i)
- (ii) Density of monohydric alcohol is less than  $H_2O$ .
- (iii) Density  $\infty$  mol. wt. (for monohydric alcohol).
- $\textbf{Solubility}: C_1 \text{ to } C_3 \text{ and t-butyl alcohol is completely soluble in } H_2O \text{ due to } H\text{-bonding}.$ (iv)

solubility 
$$\infty$$
 No. of side chainsc  $\infty$   $\frac{1}{\text{molecular weight}}$ 

#### Order of solubility:

$$C_4H_9OH \rightarrow C_5H_{11}OH \rightarrow C_6H_{13}OH$$

Number of —OH increases, H-bonding increases

(v) **Boiling points**: B.P. ∞ molecular weight

If molecular wt. is same then  $B.P. \propto \frac{1}{branching}$ 

[Number of OH increases, H-bonding increases]

- Ex. Boiling point of alcohol is more than corresponding ether. Why?
- **Sol. Reason**: H-bonding in alcohol.

Ex. Boiling point of alcohol is less than corresponding carboxylic acid. Why?

Sol. Reason: Dimer formation in carboxylic acid.

$$R-C$$
 $OH$ 
 $OH$ 
 $OH$ 

☐ Chemical properties :

Monohydric alcohol show following reactions

- (A) Reaction involving cleavage of O+H
- (B) Reaction involving cleavage of C + OH
- (C) Reaction involving complete molecule of alcohol

(A) Reaction involving cleavage of O+H: Reactivity order (Acidic nature) is

$$\mathsf{CH_3} - \mathsf{OH} \quad > \quad \mathsf{CH_3} \mathsf{CH_2} - \mathsf{OH} \quad > \quad (\mathsf{CH_3})_2 \mathsf{CH} - \mathsf{OH} \quad > \quad (\mathsf{CH_3})_3 \mathsf{C} - \mathsf{OH}$$

(i) Acidic nature :

$$H_2O > R-OH > CH \equiv CH > NH_3$$
 (Acidic strength)

Alcohols are less acidic than  $\rm H_2O$  and neutral for litmus paper and gives  $\rm H_2$  with active metals (Na, K)

$$R-OH + Na \longrightarrow R-ONa + \frac{1}{2} H_2$$

$$R-OH + K \longrightarrow R-OK + \frac{1}{2}H_2$$

(ii) Reaction with CS<sub>2</sub>:

$$R-OH + Na \longrightarrow R-ONa + \frac{1}{2}H_2$$
 $R-ONa + S=C=S \longrightarrow R-O-C-S-Na$ 
 $S$ 

Sodium alkyl xanthate (Used as floating agent)

(iii) Alkylation:

(iv) Acylation:

$$R-OH+CI-C-R \longrightarrow R-O-C-R$$

$$0 \qquad O$$

$$(Acylation)$$

$$R-OH+CI-C-CH_3 \longrightarrow R-O-C-CH_3$$

$$0 \qquad (Acetylation)$$

$$OH \qquad O-C-R$$

$$CH_3COCL \qquad O-C-R$$

Salicylic acid

Acetoxy benzoic acid Acetyl salicylic acid Aspirin [Used as analgesic] (v) Benzoylation: (Schotten Baumann's Reaction):

(Benzoylation

(vi) Esterification: Conc. H<sub>2</sub>SO<sub>4</sub> is used as catalyst and dehydrating agent.

$$R-C-OH+R-OH \xrightarrow{\text{conc. } H_2SO_4} R-C-OR+H_2O$$

$$0$$

♦ Mechanism :

$$H_2SO_4 \longrightarrow H^+ + HSO_4$$

Note: This is a laboratory method to prepare ester.

Dry HCl can be used as dehydrating agent.

- (i) Reactivity for esterification  $\propto \frac{1}{\text{Steric hinderence}}$
- (ii) Reactivity of R OH [If acid is same] :  $\mathrm{CH_3}$   $\mathrm{OH}$  > 1 > 2 > 3 alcohol
- (iii) Reactivity of RCOOH [If alcohol is same]:

(vii) Reaction with CH≡CH:

$$\label{eq:charge_charge} \begin{split} \text{CH}\!\equiv\!\text{CH} + 2\text{CH}_3 \!\!-\!\text{OH} & \xrightarrow{BF_3/\text{HgO}} & \text{CH}_3\text{CH}_3 \\ & \text{Methylal} \\ \\ \text{CH}\!\equiv\!\text{CH} + 2\text{CH}_3\text{CH}_2 \!\!-\!\text{OH} & \xrightarrow{BF_3/\text{HgO}} & \text{CH}_3\text{CH}_3 \\ & \xrightarrow{CC_2\text{H}_3} \end{split}$$

Ethylal

(viii) Reaction with carbonyl compounds :

$$R-CHO + 2R-OH \xrightarrow{H^{\oplus}} R-CH \xrightarrow{OR} CR$$

$$Acetal$$

$$R-C-R+2R-OH \xrightarrow{H^{\oplus}} R \xrightarrow{COR} CR$$

$$CH_3CHO + 2CH_3-OH \xrightarrow{H^{\oplus}} CH_3CHO$$

$$CH_3CHO + 2CH_3-OH \xrightarrow{H^{\oplus}} CH_3CHO$$

Methylal

(ix) Reaction with Grignard reagent :

$$R-Mg-X + H-OR \xrightarrow{H^{\oplus}} R-H+Mg \stackrel{X}{\bigcirc} R$$

(x) Reaction with Ketene: Ketene is used as acetylating agent.

$$CH_{2}=C=O+R-OH\longrightarrow CH_{2}=C-OH \iff CH_{3}-C=O$$

$$CC_{2}H_{5} \qquad CC_{2}H_{5}$$

$$CH_{2}=C=O+C_{2}H_{5}-OH\longrightarrow CH_{2}=C-OH \iff CH_{3}-C=O$$
Ethulacotat

(xi) Reaction with isocyanic acid: Ethyl urethane is used in preparation of urea

Ethyl urethane

(xii) Reaction with oxirane:

$$R-OH+CH_2-CH_2\xrightarrow{H^+}CH_2-CH_2$$

$$OH$$

(B) Reaction involving cleavage of C+OH: Reactivity order or basic nature is

$$\mathsf{CH_3} - \mathsf{OH} \qquad \qquad \mathsf{CH_3} \mathsf{CH_2} - \mathsf{OH} \qquad \mathsf{<} \qquad \mathsf{(CH_3)_2} \mathsf{CH} - \mathsf{OH} \qquad \mathsf{<} \qquad \mathsf{(CH_3)_3} \quad \mathsf{C} - \mathsf{OH}$$

(i) Reaction with halogen acid:

(ii) Reaction with inorganic acids:

(iii) Reaction with phosphorous halides :

(iv) Reaction with thionyl chloride (SOCl<sub>2</sub>):

(v) Reaction with  $NH_3$ : Alumina  $(Al_2O_3)$  is used as dehydrating agent.

$$R-OH+HNH_2 \xrightarrow{Al_2O_3} R-NH_2 + H_2O$$

(vi) Reaction with halogens: Oxidation and chlorination takes place simultaneously.

- (C) Reaction involving complete molecule of alcohol:
  - (i) Dehydration: Removal of H<sub>2</sub>O by two type
    - (a) Intermolecularly removal of H<sub>2</sub>O [form ether]
    - (b) Intramolecularly removal of H<sub>2</sub>O [form alkene]

$$C_2H_5OH + H_2SO_4 \xrightarrow{100 \text{ C}} (C_2H_3)_2SO_4$$

$$(conc.) \xrightarrow{140 \text{ C}} C_2H_5HSO_4$$

$$C_2H_5 - O - C_2H_5 \text{ (Williamson's synthesis)}$$

$$C_2H_2 - O - C_2H_5 \text{ (Birmination)}$$

$$C_2H_3OH + Al_2O_3$$
 $(Alumina)$ 
 $250 C$ 
 $C_2H_3-O-C_2H_3$ 
 $350 C$ 
 $C_3H_3-O-C_2H_3$ 

Ease of dehydration follow the order : 3 ROH > 2 ROH > 1 ROH >  $CH_3OH$ 

(ii) Catalytic Dehydrogenation: This reaction is useful in distinction of 1, 2 and 3 alcohols.

(iii) Oxidation: This reaction is useful in distinction of 1, 2 and 3 alcohols.

$$R-CH_{2}-OH \xrightarrow{H^{\oplus}/K_{2}Cr_{2}O_{7}} RCHO \xrightarrow{[O]} RCOOH$$

$$(p-alcohol) \qquad (same carbon acid)$$

$$R-CH-R' \xrightarrow{H/KMnO_{4} \text{ or} \atop H^{\oplus}/K_{2}Cr_{2}O_{7}} R-C-R \xrightarrow{[O]} No \text{ reaction}$$

$$OH \qquad (s-alcochol) \qquad (same carbon)$$

$$R = C-R \xrightarrow{[O]} No \text{ reaction}$$

$$R = C-R \xrightarrow{[O]}$$

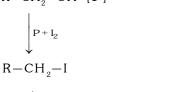
Carbonyl group goes with smaller alkyl group

(iv) Reaction with phosporous pentasulphide:

(v) Reaction with salts:

- (vi) Distinction between 1, 2 and 3 alcohols:
- (a) Lucas test: A mixture of HCl(conc.) and anhydrous ZnCl<sub>2</sub> is called Lucas reagent.

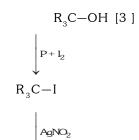




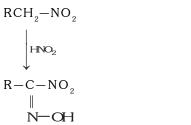


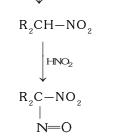
 $AgNO_2$ 

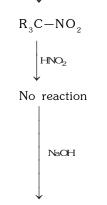
R<sub>2</sub>CH-OH [2]











Soluble (Red)

Insoluble (Blue)

NaOH

Colourless (White)

#### (vii) Dichromate test :

1 Alcohol 
$$\xrightarrow{H^{\oplus}/K_2Cr_2O_7} Acid + Cr^{+3}$$

[green]

2 Alcohol 
$$\xrightarrow{H^{\oplus}/K_2Cr_2O_7}$$
 Ketone +  $Cr^{+3}$  [green]

$$3 \ \ \, \text{Alcohol} \quad \xrightarrow{\quad H^{\oplus}/K_2Cr_2O_7 \quad \quad } \quad \text{No oxidation, No green}$$

#### (viii) Test of alcholic group:

$$R-OH \xrightarrow{Na} R-ONa + \frac{1}{2}H_{2}$$

$$[effervesence of H_{2}]$$

$$R-OH \xrightarrow{PCl_{5}} R-Cl + POCl_{3} + HCl \xrightarrow{NH_{3}} NH_{4}Cl$$

$$[White fumes]$$

R-OH Cerric ammonium nitrate Red colour

# (ix) Distinction between $CH_3$ - OH and $C_2H_5OH$

	CH <sub>3</sub> OH	CH <sub>3</sub> CH <sub>2</sub> OH
B.P.	65 C	78 C
I <sub>2</sub> + NaOH	No ppt	Yellow ppt of CHI <sub>3</sub>
Cu/300 C	Smell of formalin [HCHO]	No smell
Salicylic acid	Smell of oil of wintergreen	No smell

#### ■ Additional reactions :

### (a) Oxidation by $\mathrm{HIO}_4$ [per iodic acid] :

#### ◆ Condition for oxidation by HIO<sub>4</sub>:

At least 2 -OH or 2 >C=O or 1 -OH and 1 >C=O should be at adjacent carbons.

Example : 
$$CH_3$$
  $CH_3$   $CH_4$   $CH_5$   $CH_5$ 

#### Example:

#### (b) Pinacole - Pinacolone Rearrangement :

Pinacole Pinacolone

#### Mechanism:

# AROMATIC HYDROXY DERIVATIVES

#### ☐ Phenolic compounds :

Compounds in which -OH group is directly attached to sp<sup>2</sup>c [Benzene ring]

All phenolic compounds give characteristic colour with neutral FeCl<sub>2</sub>.

$$\begin{array}{ccc} \text{Ph-OH} & \xrightarrow{\text{neutral FeCl}_3} & \text{Violet colour} \\ \\ \text{CH}_3\text{CH}_2\text{-OH} & \xrightarrow{\text{neutral FeCl}_3} & \text{No colour} \end{array}$$

# PHENOL (C<sub>6</sub>H<sub>5</sub>OH)

Phenol is also known as carbolic acid or Benzenol or hydroxy benzene. In phenol -OH group is attached with  $sp^2$  hybridised carbon. It was discovered by Runge in the middle oil fraction of coaltar distillation and named it carbolic acid (carbo = coal; oleum = oil) . It is also present in traces in human urine.

#### ☐ General Methods of preparation :

#### (1) From benzene sulphonic acid:

When sodium salt of benzene sulphonic acid is fused with NaOH phenol is obtained.

$$C_6H_5SO_3Na + NaOH \longrightarrow C_6H_5OH + Na_2SO_3$$

#### (2) From benzene diazonium chloride:

When benzene diazonium chloride solution is warmed, phenol is obtained with evolution of nitrogen.

$$\begin{array}{c}
N_2CI \\
& OH \\
\hline
O \\
& A
\end{array}$$

$$\begin{array}{c}
\text{(Steam distilled)} H_2O \\
\hline
O \\
& A
\end{array}$$

$$\begin{array}{c}
+ N_2 + HCI
\end{array}$$

(3) By distilling a phenolic acid with sodalime (decarboxylation):

$$\begin{array}{ccc}
OH & OH \\
& & OH \\
&$$

Salicylic acid

(4) From Grignard reagent: (The Grignard reagent on reaction with oxygen and subsequent hydrolysis by acid yields phenol)

$$C_6H_5MgBr \xrightarrow{[O]} C_6H_5OMgBr \xrightarrow{H_2O} C_6H_5OH + Mg OH$$

(5) From benzene:

$$\bigcirc + [O] \qquad \xrightarrow{V_2O_5} \qquad \bigcirc \bigcirc$$

(6) From chloro benzene:

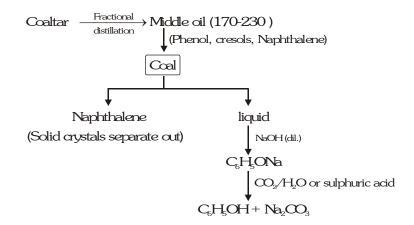
Order of NSR:

#### (7) Industrial preparation of phenol:

Phenol can be prepared commercially by:

- (a) Middle oil fraction of coaltar distillation
- (b) Cumene
- (c) Raschig process
- (d) Dow's process

(a) Middle oil fraction of coaltar:



(b) From cumene (Isopropyl benzene): Cumene is oxidised with oxygen into cumene hydroperoxide in presence of a catalyst. This is decomposed by dil.  $H_2SO_4$  into phenol and acetone.

(c) Raschig process: Chlorobenzene is formed by the interaction of benzene, HCl and air at  $300^{\circ}$  C in presence of catalyst CuCl<sub>2</sub> + FeCl<sub>3</sub>. It is hydrolysed by superheated steam at  $425^{\circ}$  C to form phenol and HCl

$$\begin{array}{cccc} C_6H_6 + HCl + \frac{1}{2}O_2 & \xrightarrow{CuCl_2/FeCl_3} & C_6H_5Cl + H_2O \\ \hline \\ C_6H_5Cl + H_2O & \xrightarrow{425^0C} & C_6H_5OH + HCl \\ \text{(super heated steam)} \end{array}$$

(d) Dow process: This process involves alkaline hydrolysis of chloro benzene-(large quantities of phenol formed).

$$C_6H_5Cl + NaOH \xrightarrow{Cu-Fe} OH + NaCl$$

### ■ Physical properties :

- (i) Phenol is a colourless, hygroscopic crystalline solid.
- (ii) It attains pink colour on exposure to air and light. (slow oxidation)

Phenoquinone(pink colour)

- (iii) It is poisonous in nature but acts as antiseptic and disinfectant.
- (iv) Phenol is slightly soluble in water, readily soluble in organic solvents.
- (v) Solublity of phenol in water is much lower than alcohols because of larger hydrocarbon part in the molecule.
- (vi) Due to intermolecular H-Bonding, phenol has relatively high boiling point than the corresponding hydrocarbons, aryl halides etc. but intermolecular H-bonding in o-derivatives is used in the preparation of dyes, drugs, bakelite and it's melting point (MP) is 43 C and boiling point (BP) is 182 C.

#### ☐ Chemical Properties :

#### (A) Reactions due to -OH group:

◆ Acidic Nature: Phenol is a weak acid. The acidic nature of phenol due to formation of stable phenoxide ion in solution. The phenoxide ion is stable due to resonance. The negative charge is spread through out the benzene ring which is stabilising factor in the phenoxide ion. Electron withdrawing groups (-NO<sub>2</sub>, -Cl) increase the acidity of phenol while electron releasing groups (-CH<sub>3</sub> etc.) decrease the acidity of phenol.

$$C_6H_5OH$$
  $\longrightarrow$   $C_6H_5O + H_3O$ 

Phenol is stronger acid than alcohols but weaker than the carboxylic acids and even carbonic acid. The acidic nature of phenol is observed in the following:

- (i) Phenol changes blue litmas to red.
- (ii) Highly electro positive metals react with phenol.

$$2C_6H_5OH + 2Na \longrightarrow 2C_6H_5ONa + H_9$$

(iii) Phenol reacts with strong alkalies to form phenoxides.

$$C_6H_5OH + NaOH \longrightarrow C_6H_5ONa + H_2O$$

(iv) However phenol does not decompose  $Na_2CO_3$  or  $NaHCO_3$  because phenol is weaker than carbonic acid.

(v) Phenol does not react with NaHCO<sub>3</sub>.

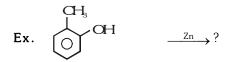
- (vi) Acetic acid reacts with  $\mathrm{NaHCO}_3$  and gives effervesence of  $\mathrm{CO}_2.$
- Reaction with PCl<sub>5</sub>: Phenol reacts with PCl<sub>5</sub> to form chloro benzene. The yield of chlorobenzene is poor and mainly triphenyl phosphate is formed.

$$C_6H_5OH + PCl_5 \longrightarrow C_6H_5Cl + POCl_3 + HCl$$
  
 $3C_6H_5OH + POCl_3 \longrightarrow (C_6H_5)_3PO_4 + 3HCl$ 

• Reaction with Zn dust: When phenol is distilled with zinc dust benzene is obtained.

$$C_6H_5OH + Zn \longrightarrow C_6H_6 + ZnO$$

Sol. No reaction



$$\qquad \qquad \text{Sol.} \quad \bigcirc$$

Ex. 
$$CH_3$$
  $Z_n \to 3$ 

Sol. No reaction

Ex. 
$$OH$$
  $\xrightarrow{Zn}$ 

♠ Reaction with NH<sub>3</sub>( Bucherer reaction): Phenol reacts with NH<sub>3</sub> in presence of anhydrous ZnCl<sub>2</sub> to form aniline.

• Reaction with FeCl<sub>3</sub>: Phenol gives violet colouration with FeCl<sub>3</sub> solution (neutral) due to formation of a complex.

$$C_6H_5OH + FeCl_3 \longrightarrow Voilet colour$$

This reaction is used to differentiate phenol from alcohols.

 Acetylation (Schotten-Baumann reaction): Phenol reacts with acid chlorides or acid anhydrides in alkali solution to form phenyl esters.

◆ Ether formation (Alkylation): Phenol reacts with alkyl halides in alkali solution to form phenyl ethers. (Williamson's synthesis)

- (B) Reaction of Benzene Ring: The −OH group is ortho and para directing. It activates the benzene nucleus.
  - lack Halogenation: Phenol reacts with bromine in  $CCl_4$  to form mixture of o-and p-bromo phenol.

$$\begin{array}{c}
OH \\
OH \\
\hline
O + Br_2 \xrightarrow{CHCl_3 \text{ or } CS_2 \text{ or } CHCl_4}
\end{array}$$

$$\begin{array}{c}
OH \\
\hline
OH \\
\hline
OH
\end{array}$$

$$\begin{array}{c}
OH \\
\hline
OH
\end{array}$$

$$\begin{array}{c}
OH \\
\hline
OH$$

$$\begin{array}{c}
OH \\
\hline
OH
\end{array}$$

Phenol reacts with bromine water to form a white ppt. of 2,4,6 tribromo phenol.

$$OH \longrightarrow Br \longrightarrow Br \longrightarrow Br + 3HBr$$

lacktriangle Nitration: Phenol reacts with dil. HNO $_3$  at 0-10 C to form o- and p- nitro phenols.

$$\begin{array}{cccc}
OH & OH & OH \\
& & & \\
O & & & \\
OH & OH \\
NO_2 & + & OH \\
& & & \\
OH & & \\
NO_2 & + & OH \\
& & & \\
NO_3 & + & OH \\
& & & \\
NO_4 & + & OH \\
& & & \\
NO_5 & + & OH \\
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When phenol is treated with nitrating mixture to form 2,4,6- trinitro phenol (picric acid)

 Sulphonation: Phenol reacts with fuming H<sub>2</sub>SO<sub>4</sub> to form o-and p-hydydroxy benzene sulphonic acid at different temperatures.

Friedel-Craft's reaction: Phenol when treated with methyl chloride in presence of anhydrous AlCl<sub>3</sub> p-cresol is main product.

 Gattermann aldehyde synthesis: When phenol is treated with liquid HCN and HCl gas in presence of anhydrous AlCl<sub>3</sub> yields mainly p- hydroxy benzaldehyde (formylation)

$$\begin{array}{c}
OH \\
OH \\
OH
\end{array}$$

$$\begin{array}{c}
OH \\
H_2O \\
-NH_3
\end{array}$$

$$\begin{array}{c}
OH \\
OH
\end{array}$$

 $\underline{\hspace{1cm}}^{AlCl_3}$  HN = CHCl

HCl + HCN

• Riemer-Tiemann reaction: Phenol on refluxing with chloroform and NaOH (aqueous) followed by acid hydrolysis yields o-hydroxy benzaldehyde. When CCl<sub>4</sub> is used salicylic acid is formed.

**Mechanism** :  $CCl_2$  is neutral attacking electrophile (formed by  $\alpha, \alpha-$  elimination reaction)

★ Kolbe 's Schmidt reaction: This involves the reaction of C<sub>6</sub>H<sub>5</sub>ONa with CO<sub>2</sub> at 140<sup>0</sup> C followed by acid hydrolysis salicylic acid is formed followed.

◆ Hydrogenation: Phenol when hydrogenated in presence of Ni at 150-200°C forms cyclohexanol.

$$\begin{array}{c}
OH \\
O \\
\hline
O \\
+ 3H_2 \\
\hline
150-200 C
\end{array}$$
Cyclohexanol. (C<sub>6</sub>H<sub>11</sub>OH)

(used as a good solvent)

♦ Fries rearrangement reaction :

$${\rm C_6H_5OH} \ + \ {\rm CH_3COCl} \ \ \underline{\quad \quad } \ {\rm ^{Anhydrous} \ AlCl_3} \ \longrightarrow \ {\rm C_6H_5OCOCH_3}$$

◆ Duff 's reaction: This method gives only the o-compound which is hindered by the presence of a -I group in the ring.

 Coupling reactions: Phenol couples with benzene diazonium chloride in presence of an alkaline solution to form a dye (p- hydroxy azobenzene) red only.

$$\begin{array}{c|c}
\hline
\bigcirc N_2O + & \bigcirc OH \xrightarrow{NaOH} & \bigcirc N=N-\bigcirc O-OH \\
\hline
& p-hydroxy azobenzene (Orange dye or Red dye)
\end{array}$$

Phenol couples with phthalic anhydride in presence of conc.  $\rm H_2SO_4$  to form a dye (phenolphthalien) used as an indicator.

Phenol (2 molecules)

Phenolphthalien (Colourless in acidic medium and pink in alkaline medium)

lacktriangle Lederer Manasse (Condensation with formaldehyde): Phenol condenses with HCHO (excess) in presence of NaOH or weak acid ( $H^+$ ) to form a polymer known as bakelite (aresin).

Polymer bakelite (Phenol formaldehyde resin)

 $\textbf{Leibermann's nitroso reaction}: \ When \ phenol \ is \ reacted \ with \ NaNO_2 \ and \ conc. \ H_2SO_4 \ it \ gives \ a$ deep green or blue colour which changes to red on dilution with water. When made alkaline with NaOH original green or blue colour is restrored.

This reaction is used as a test of phenol.

$$2\text{NaNO}_2 + \text{H}_2\text{SO}_4 \longrightarrow 2\text{HNO}_2 + \text{Na}_2\text{SO}_4$$

Reaction with acetone: (Condensation with acetone)

Oxidation:

1, 4 – Dihydroxy benzene

### ☐ Test of Phenol:

- (i) Phenol turns blue litmus to red.
- (ii) Ageous solution of phenol gives a violet colour with a drop of ferric chloride.
- (iii) Phenol gives Lieber mann 's nitroso test.

- (iv) Ageous solution of phenol gives a white ppt. of 2,4,6 tribromophenol with bromine water.
- (v) Phenol combines with phthalic anhydride in presence of conc.  $H_2SO_4$  to form phenolphthalein which gives pink colour with alkali.
- (vi) With ammonia and sodium hypochlorite, phenol gives blue colour.

## figspace Differences between phenol and alcohol ( $C_2H_5OH$ ):

- (i) Phenol is more acidic than aliphatic alcohol due to resonance in phenoxide ion.
- (ii) Phenol gives violet colour with FeCl<sub>3</sub> while aliphatic alcohol does not give.
- (iii) Phenol gives triphenyl phosphate with PCl<sub>5</sub> while aliphatic alcohol gives alkyl chloride.
- (iv) Phenol has phenolic odour whereas alcohol has pleasent odour.
- (v) Phenol on oxidation gives quinone while alcohol gives aldehyde or ketone and acids.

#### ☐ Uses of Phenol: Phenol is used:

- (a) As an antiseptic in soaps and lotions. "Dettol" (2,4-Dichloro-3,5-dimethyl phenol)
- (b) In manufacture of azodyes, phenolphthalein, picric acid (explosive), cyclohexanol (Solvent for rubber), plastics (bakelite) etc.
- (c) In manufacture of drugs like aspirin salol, phenacetin etc.
- (d) As preservative for ink.

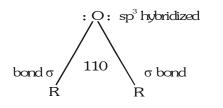
# **ETHER**

R-O-R (Dialkyl ether), alkoxy alkane. It's General formula is  $C_nH_{2n+2}O$ .  $CH_3-O-CH_2CH_3$  (Methoxy ethane) or ethyl methyl ether or 2-oxa butane Ether is monoalkyl derivative of R-OH and dialkyl derivative of H<sub>2</sub>O

Classification: They may be classified as:

- Simple or symmetrical ether. e.g. R-O-R (a)
- Mixed or unsymmetrical ether e.g. R-O-R' (b)

#### Structure:



The molecule of ether is bent due to lone pair of electron on oxygen atom-bond electron repulsion. The bond angle is  $110^{\circ}$ . It is greater than that of water  $105^{\circ}$  due to the repulsion between bulky alkyl groups. Due to bent structure, it posses dipole moment and hence are polar molecules.

- General Methods of Preparation:
- (A) From alkyl halides:
  - By Williamson's synthesis : (i)

$$R{-}X \ + \ Na{-}O{-}R \qquad \longrightarrow \qquad R{-}O{-}R \ + \ NaX \ [\, {\boldsymbol{S}}_{N^2} \ \ Reaction]$$

**Mechanism** :  $[S_{N^2}]$  Reaction]

$$C_2H_3ON_a^{\oplus} \rightleftharpoons C_2H_3O^{\ominus} + N_a^{\oplus}$$

$$C_{2}H_{5}\overset{\Theta}{\longrightarrow}C_{-}I \xrightarrow{Slow} C_{2}H_{5}O - C - I \xrightarrow{Fast} C_{2}H_{5}-O-CH_{3}+I^{\Theta}$$

$$N_a^{\oplus} + I^{\ominus} \longrightarrow N_a I$$

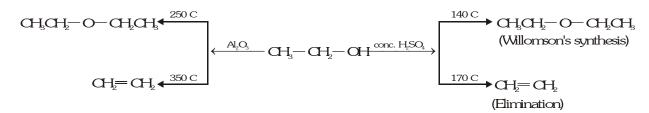
$$\begin{array}{c} \text{Example} \ : \ \begin{array}{c} CH_3 \\ -C-C+C+CH_3ONa \longrightarrow H_3C-C \\ -CH_3 \end{array} \end{array} \xrightarrow{CH_3} \begin{array}{c} CH_2 \\ -C-C \\ -CH_3 \end{array}$$

### (ii) Reaction with Dry Ag<sub>2</sub>O:

$$2RX + Ag_2O$$
  $\xrightarrow{\Delta}$   $R-O-R + 2AgX$ 

#### (B) From R-OH:

(i) By dehydration : 
$$R-OH \xrightarrow{Con. H_2SO_4 \land} ?$$



#### (ii) Reaction with $CH_2N_2$ (diazomethane):

$$R-OH + CH_2-N_2 \xrightarrow{\Delta} R-O-CH_2-H + N_2$$

#### ☐ Physical Properties :

- (i)  $CH_3OCH_3$ ,  $CH_3OCH_2CH_3$  are gases and higher are volatile liquids.
- (ii) Ether are less polar  $[\mu=1.18D]$ .
- (iii) Ethers are less soluble in  $H_2O$ .
- (iv) Ethers have less BP then corresponding alcohol.
- Ex. Ethers are less soluble in  $H_2O$ . Why?
- Sol. Reason: Due to less polar, it forms weaker H-Bonding with H<sub>2</sub>O.
- Ex. Ethers have less BP then corresponding alcohol. Why?
- Sol. Reason: No H-Bonding in ether molecules.

### ☐ Chemical properties :

Ethers are less polar so less reactive and do not react with active metals [Na,K], cold dil. acid, oxidising and reducing agent.

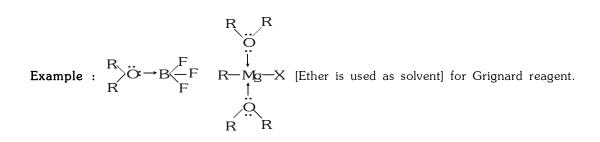
Reason: They do not have any active functional group.

# 1. Basic nature : Due to presence of $\ell.p$ on oxygen atom ether behave as lewis base

Ethers react with cold conc. acid and form oxonium ion

Example: 
$$C_2H_5 = \overset{\circ}{\square} - C_2H_5 = \overset{\text{cold}; \text{ conc.}}{H_2SO_4} + \begin{bmatrix} C_2H_5 = \overset{\oplus}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\ H \end{bmatrix} + \begin{bmatrix} C_2H_5 = \overset{\circ}{\bigcirc} - C_2H_5 \\$$

Ether form dative bond with Lewis acids like BF<sub>3</sub>, AlCl<sub>3</sub>, RMgX etc.



2. Halogenation:

$$\begin{array}{c|c} & C_2/\operatorname{dark} & C_1 & C_2 & C_3 & C_4 & C_4 & C_5 &$$

3. Formation of peroxides: Ether add up atmospheric oxygen or ozonised oxygen. It is explained by Free radical mechanism as intermediates is free radical.

$$C_{2}H_{5}-O-C_{2}H_{5} \xrightarrow{O_{2}(\text{nonpolar})} CH_{3}CH_{2}-O-CHCH_{3}$$

$$(\text{Non polar}) \quad \text{sunlight or UV} \qquad O-O-H$$

$$C_{2}H_{5}\overset{\bullet}{\square}C_{2}H_{5}+\overset{\bullet}{\square}\longrightarrow C_{2}H_{5}\overset{\bullet}{\square}C_{2}H_{5} \text{ or } (C_{2}H_{3})_{2}O\longrightarrow O$$

$$\vdots\overset{\bullet}{\square}c$$

$$CH_{3}CH_{2}-O-CH_{2}-Ph \xrightarrow{O_{2}} CH_{3}-CH_{2}-O-\overset{\bullet}{C}H-Ph \xrightarrow{O_{2}} CH_{3}-CH_{2}-O-\overset{\bullet}{C}H-Ph$$

$$\text{stable by resonance}$$

$$\text{Peroxides are unstable and explosives.}$$

Peroxides are unstable and explosives.

Test for peroxides ether (peroxide) 
$$\xrightarrow{FeSO_4/KCNS}$$
 Red colour ether (Peroxides) +  $Fe^{+2}$   $\xrightarrow{Fe^{+3}}$   $\xrightarrow{CNS}$   $Fe(CNS)_3$  (Red)

- Reaction with hot dil.  $H_2SO_4$ : R-O-R  $\xrightarrow{\text{hot dil}}$  2R-OH
- 5.

- 6.
- $\textbf{Reaction with BCl}_{\textbf{3}} \ : \qquad \qquad 3\text{ROR} \ + \ \text{BCl}_{3} \qquad \longrightarrow \qquad 3\text{RCl} \ + \ (\text{RO})_{3}\text{B}$ 7.

8. Reaction with RCOCl : ROR + RCOCl  $\xrightarrow{AlCl_3}$  RCOOR +RCl

9. Reaction with CO: ROR + CO  $\xrightarrow{BF_3/HgO 500 \text{ atm}}$  RCOOR

10. Reaction with  $C_2H_5Na$ :  $CH_2CH_2-O-CH_2CH_2+C_2H_5 \atop H$   $CH_3CH_2OH+CH_2=CH_2+C_2H_6 \atop H$  Stronger

11. Dehydration :  $CH_3CH_2-C-CH_2CH_3 \xrightarrow{Al_2O_3} 2CH_3=CH_3+H_3C$ 

 $\textbf{12.} \quad \textbf{Reduction} \quad : \qquad \qquad \text{CH}_{3}\text{CH}_{2}\text{OCH}_{2}\text{CH}_{3} \qquad \xrightarrow{\begin{array}{c} \text{RedP+HI} \\ \text{heat} \end{array}} \qquad 2\text{CH}_{3}\text{CH}_{3}$ 

13. Oxidation:

 $\text{CH}_3\text{CH}_2-\text{O}-\text{CH}_2\text{CH}_3 \xrightarrow{\text{$H^+/\text{K}_2\text{Cr}_2\text{O}_7$}} 2\text{CH}_3\text{CH}_2\text{OH} \xrightarrow{\text{[O]}} 2\text{CH}_3\text{CHO} \xrightarrow{\text{[O]}} 2\text{CH}_3\text{COOH}$ 

**14.** Combustion :  $C_2H_5OC_2H_5 + 6O_2 \longrightarrow 4CO_2 + 5H_2O$  (explosive mixture)

15. Reaction with HX : Reactivity of HX | HI>HBr>HCl

Reaction with cold conc. HX:

Ethers forms oxonium salt with cold and conc. HCl (less reactive) Cold conc. HI and HBr (more reactive) break C-O bond.

Ex. 
$$CH_3$$
  $C-C-C-CH_2-CH_3$   $Cold and conc.$   $COld and$ 

Sol. Mechanism

 $\bullet \qquad \text{If oxonium ion gives more stable carbocation } [Ph\overset{\oplus}{CH_2}, CH_2 = CH - \overset{\oplus}{C}H_2, \ (CH_3)_3\overset{\oplus}{C}] \text{ then } SN^1 \text{ reaction occurs.}$ 

 $\bullet \qquad \text{If oxonium ion gives less stable carbocation } [P\stackrel{\oplus}{h}\,,\, CH_2 \stackrel{\oplus}{=} \stackrel{\oplus}{C}\,H,\,\, CH_3\stackrel{\oplus}{C}\,H_2] \text{ then } SN^2 \text{ reaction occurs,} \\ \text{and } X^{\Theta} \text{ attacks at less hindered carbon.}$ 

**Ex.**  $CH_3CH_2-O-CH_2Ph \xrightarrow{Cold\ con.} CH_3CH_2-OH + PhCH_2-I$ , write mechanism of given reaction.

Sol. Mechanism:

$$\begin{array}{c} \text{CH}_3\text{CH}_2\overset{\cdots}{\text{CH}_2}\text{Ph} \overset{\text{H}^\oplus}{\longleftarrow} \text{CH}_3\text{CH}_2 & \xrightarrow{\text{I}^\oplus} \text{PhCH}_2\text{I} + \text{CH}_3\text{CH}_2 & \xrightarrow{\text{I}^\oplus} \text{PhCH}_2\text{I} + \text{CH}_3\text{CH}_2 & \xrightarrow{\text{I}^\oplus} \text{PhCH}_2\text{I} \\ \text{H} \end{array}$$

**Ex.** 
$$CH_3CH_2-O-CH_3 \xrightarrow{anhy.HI} ?$$

Sol. 
$$CH_{2}CH_{2}-\overset{\circ}{C}-CH_{3}\overset{H}{\Longrightarrow}CH_{2}CH_{2}-\overset{\circ}{C}-CH_{3}\overset{|0}{\Longrightarrow}CH_{3}I+CH_{2}CH_{2}CH_{3}$$

Oxonium ion gives less stable carbocation  ${\rm SN}^2 \ {\rm reaction} \ {\rm I}^\Theta \ {\rm attacks} \ {\rm at \ less \ hinderd \ carbon}.$ 

Ex. 
$$CH_3 - CH_2 - O - Ph \xrightarrow{(aq.)HBr}$$
 ?

- If excess of HI is used then two moles of alkyl hallides are formed.
- (B) Reaction with hot and conc. HX:

$$\mathsf{CH_3CH_2} - \mathsf{O} - \mathsf{CH_3} \quad \xrightarrow{\quad \mathsf{hot} \ \mathsf{and} \ \mathsf{conc} \ \mathsf{HI} \quad} \quad \mathsf{CH_3CH_2} - \mathsf{I} \ + \ \mathsf{CH_3} - \mathsf{I}$$

Ex. 
$$C_2H_5-O-C_2H_5 \xrightarrow{\text{hot and conc. HBr}}$$
 ? + ?

**Sol.** 
$$C_2H_5$$
-Br +  $C_2H_5$  - Br

#### ☐ Uses of ether:

- (i) General anaesthetics agent.
- (ii) Refrigerant a mixture of ether and dry ice gives temperature as low 110 C.
- (iii) Solvent for oil, fats, resins, Grignard reagent.
- (iv) For providing inert & moist free medium to organic reaction example: Wurtz reactions.
- (v) In perfumery.
- (vii) Mixture of alcohol and ether is used as a substitute of petrol. Trade name "Natalite"
- (viii) Halothane ( $CF_3CHClBr$ ) used as an anaesthetic because it produces unconsciousness without affecting lung and heat.

#### $\Box$ Preparation of Epoxides:

- (i) Epoxidation of alkenes by reaction with peroxy acids
- (ii) Base-promoted ring closure of vicinal halohydrins
- (iii) Epoxidation of alkenes by reaction with peroxy acids

♦ Epoxidation of alkenes by reaction with peroxy acids :

$$C=C + R-C-OOH \longrightarrow C-C + R-C-OH$$
Peroxy acid Epoxide

Example:

(a) 
$$CH_2 = CH_1(CH_2)_0 - CH_3 + CH_3 - CC - COOCH \longrightarrow CH_2 - CH_1(CH_2)_0 CH_3 + CH_3 - CC - COCH$$

(c) Epoxidation is a stereospecific syn addition :

$$C = C + C^{H} + CH^{2} - C - COH$$

$$C^{2}H^{2} + CH^{2} - C - COH$$

$$C^{2}H^{2} + CH^{2} - C - COH$$

$$C^{2}H^{2} + CH^{2} - C - COH$$

(E) -1,2-diphenyl ethene

trans -2,3-diphenyl oxirane

#### Mechanism:

♦ Base-promoted ring closure of vicinal halohydrins :

$$R_{2}C = CR_{2} \xrightarrow{\frac{x_{2}}{H_{2}o}} R_{2}C \xrightarrow{R_{2}} R_{2}C \xrightarrow{HO} R_{2}C \xrightarrow{CR_{2}}$$

$$HO X$$
Vicinal halohydrin Epoxide

Mechanism:

#### Example:

- Reaction of Epoxides :
- ♦ With Grignard reagent :

$$RMgX + H_2C \longrightarrow H_2 \qquad \xrightarrow{\text{(1) diethylether}} RCH_2CH_2OH \text{ (primary alcohol)}$$

♦ Nucleophilic ring opening reactions of epoxides :

$$Y: + R_2C \longrightarrow R_$$

 $\textbf{Note}: \qquad \text{Nucleophilic ring opening reactions of epoxides is the characteristic feature of } S_{N^2} \text{ reaction}.$ 

### ♦ Nucleophilic ring opening of epoxides :

Mechanism

Step-1: 
$$\begin{array}{c} H_{2}C - CH_{2} + H - C \\ \vdots \\ H \end{array}$$

Step-2: 
$$H \xrightarrow{H_2C - CH_2} H \xrightarrow{Slow} H \xrightarrow{+} H$$

Example :

1,2-Epoxycydohexane

trans -2-bromo cyclohexanol

$$H_3C$$
 $CH_3$ 
 $CH_3OH$ 
 $H_2SO_4$ 
 $CH_3OH$ 
 $CH_3$ 

2,2,3-trimethyl oxirane

3-methoxy-3-methyl-2-butanol