



# ENGINEERING CHEMISTRY

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Department of Science and Humanities

## *Class content:*

- *The allowed rotational energies of a rigid diatomic molecule*
- *Selection rule*
- *Rotational spectrum*

# Module I- Molecular Spectroscopy

The energy expressed in spectroscopic units( $\text{cm}^{-1}$ ) is given by :

$$\varepsilon_J = \frac{h}{8\pi^2 I c} J(J+1) \text{cm}^{-1}$$

which can be written as

$$\varepsilon_J = BJ(J+1) \text{cm}^{-1}$$

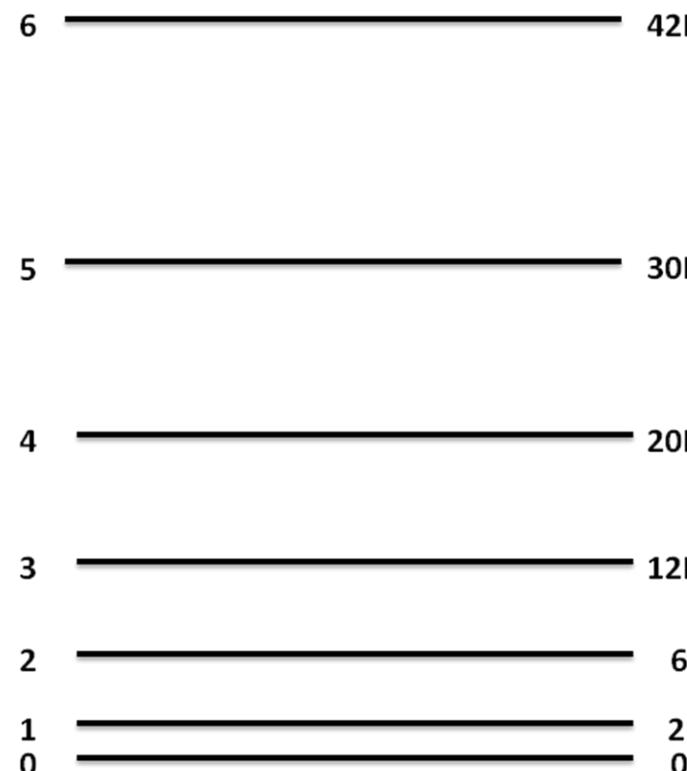
where  $B = \frac{h}{8\pi^2 I c} \text{cm}^{-1}$

B is known as the **rotational constant**

Substituting for values of

$J = 0, 1, 2, 3, \dots$ , we can get the energies  
for the rotational levels

J	$\varepsilon_J$
0	0
1	2B
2	6B
3	12B



# Module I- Molecular Spectroscopy

The **selection rules** for rigid rotor model obtained after solving Schrodinger equation is :

- **Gross selection rule – molecule should possess permanent dipole moment**
- $\Delta J = \pm 1$

Since  $\varepsilon_J = BJ(J+1)cm^{-1}$

For rotational transition of a molecule from level  $J \rightarrow J+1$  , the energy absorbed is given by  $\Delta\varepsilon_{J \rightarrow (J+1)} = \bar{\nu} = 2B(J+1)cm^{-1}$

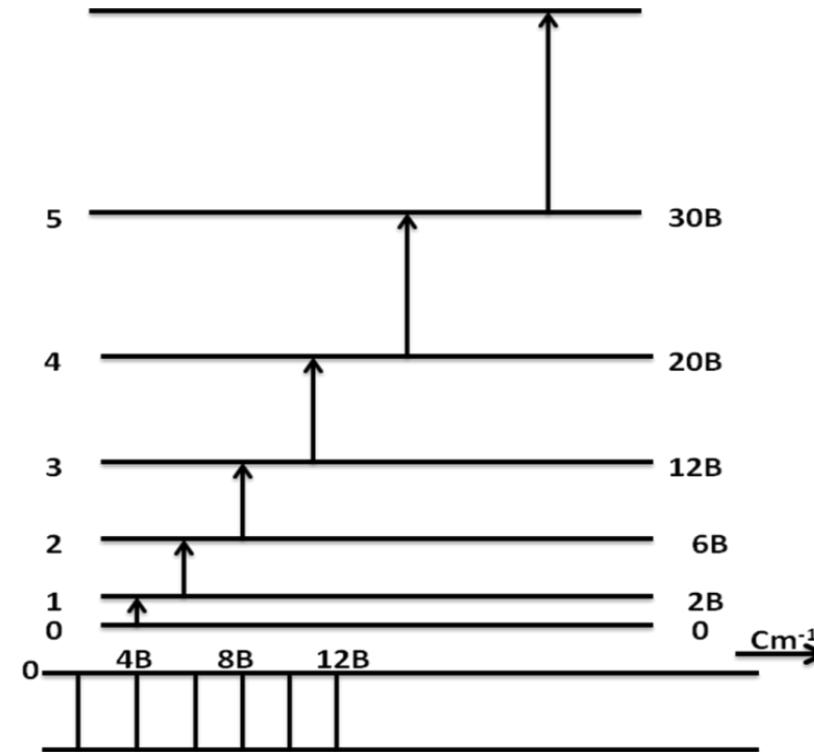
Substituting for values for  $J = 0,1,2,3....$

$J$	$\Delta\varepsilon_{(J \rightarrow J+1)}$
0	$2B\text{ cm}^{-1}$
1	$4B\text{ cm}^{-1}$
2	$6B\text{ cm}^{-1}$

# Module I- Molecular Spectroscopy

## Rotational energy levels and spectrum

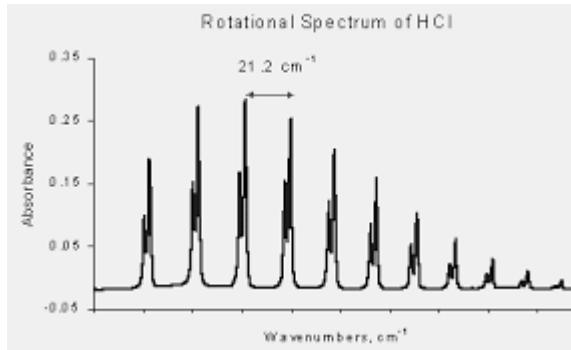
J	$\Delta\epsilon_{(J \rightarrow J+1)}$
0	$2B \text{ cm}^{-1}$
1	$4B \text{ cm}^{-1}$
2	$6B \text{ cm}^{-1}$
3	$8B \text{ cm}^{-1}$
4	$10B \text{ cm}^{-1}$



Source:Fundamentals of Molecular Spectroscopy: C. N. Banwell and Elaine M McCash, Fifth Edition, MCGRAW-HILL Education (India) Private Ltd.

# Module I- Molecular Spectroscopy

## Information obtained from the rotational spectrum



Source:<http://www.physics.dcu.ie/~be/Ps415/Rotational1.pdf>

- The **first line in the spectrum** appears at  $2B \text{ cm}^{-1}$  and the **distance between any two consecutive lines** is constant and is equal to  $2B \text{ cm}^{-1}$ . We can get value of 'B' from the spectrum and calculate  $I$ , the moment of inertia using the expression

$$B = \frac{\hbar}{8\pi^2 I c} \text{ cm}^{-1}$$

Since  $I = \mu r_o^2$ ,  $r_o$  can be determined ;  $r_o = \sqrt{\frac{I}{\mu}}$   
 $r_o$  is the **bond length** of the molecule

- The spectrum also reveals that **some higher rotational levels are also populated at room temperature**



**THANK YOU**

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