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CBSE 10th Periodic Classification Of Elements Unsolved Paper

SECTION - A

Q.1. X, Y and Z are the elements of a dobereiners triad. If the atomic mass of x is 7 and that of z is 39, what should be the atomic mass of y?

Q.2. A and B are the two elements having similar properties which obey Newlands law of octaves. How many elements are there in between A and B?

Q.3. Which of the following statements is not correct statement about the trends when going from left to right across the periods of Periodic Table?

- (a) The elements become less metallic in nature.
- (b) The number of valance electrons increases.
- (c) The atoms lose their electrons more easily.
- (d) The oxides becomes more acidic.

Q4. Element X forms a chloride with the formula XCl_2 which is a solid with a high melting point. X would most likely be in the same group of the Periodic Table as

- (a) Na
- (b) Mg
- (c) Al
- (d) Si

Q.5. Besides gallium, which other elements have since been discovered that were left by Mendeleev in his Periodic Table?

Q.6. Which of the following is the most reactive halogen?

- (a) F
- (b) Cl
- (c) Br
- (d) I

Q.7. The number of periods in the long form of the periodic table is

- (a) 6
- (b) 7
- (c) 10
- (d) 18

Q.8. The law of octaves was proposed by

- (a) Newland
- (b) Doberiener
- (c) Lothar Meyer
- (d) Mendeleev

Q.9. Which of the following has maximum atomic size?

- (a) K
- (b) Ca
- (c) Al
- (d) P

Q.10. Name three elements which behave as metalloids?

SECTION - B

Q.11. Nitrogen (atomic number 7) and phosphorus (atomic number 15) belong to group 15 of the Periodic Table. Write the electronic configuration of these two elements. Which of these will be more electronegative? Why?

Q.12. (a) What property do all elements in the same column of the Periodic Table as boron have in common?

(b) What property do all elements in the same column of the Periodic Table as Fluorine have common?

Q.19. An element X (2,82) combines separately with NO and $(\text{SO}_4)^{2-}$, (PO⁻ -radicals. Write the formulae of the three compounds so formed. To which group of the periodic table does the element 'X' belong? Will it form covalent or ionic compound? Why?

Q.20. Two elements X and Y have atomic numbers 12 and 16 respectively. Write the electronic configuration for these elements. To which period of the modern periodic table do these two elements belong? What type of bond will be formed between them and why?

Q.21. Which element has

- (a) two cells, both of which are completely filled with electrons?
- (b) the electronic configuration 2, 8, 2?
- (c) a total of three shells, with four electrons in its valance shell?
- (d) twice as many electrons in its second shell as in its first shell?

Q.22. (a) Lithium, sodium and potassium are all metals that react with water to liberate hydrogen gas. Is there any similarity in the atoms of these elements?

- (b) Helium is an un-reactive gas and neon is a gas of extremely low reactivity. What, if anything, do their atoms have in common?

Q.23. Name:

- (a) Three elements that have a single electron in their outermost shells.
- (b) Two elements that have two electrons in their outermost shells.
- (c) Three elements with filled outermost shells.

Q.24. What were the limitations of Newland's law of octaves?

SECTION – D

Q.25. Calcium is an element with $Z = 20$

- (a) Is it a metal or a non-metal?
- (b) Will its size be bigger or smaller than that of potassium?
- (c) write the formula of its chloride

Q.27. (a) Name the elements present in the third period and classify them into metals and non-metals.

(b) On which side of the table do you find the metals?

(c) On which side of the table do you find the non-metals?

Q.28. Write two major shortcomings of Mendeleev's periodic table? How have these been removed in the modern periodic table?

Q.29. Two elements X and Y belong to Group 1 and Group 2 respectively in the same period. Compare them with respect to:

(a) The number of valence electrons

(b) Valency

(c) metallic character

(d) Size of the atoms

(e) Formulae of their oxides and chlorides.

Q.30. What were the limitations of Newland's law of octaves?



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