Start

Yes

Write the name of the metal first, followed by the name of the non metal. The non metal's name should be changed to end in "ide" Remember to add the relevant roman numeral after the metal

|  |  |  |  |
| --- | --- | --- | --- |
| Element | Symbol | Ionic Charges | Roman Numeral |
| Copper | Cu | 1+,2+ | I,II |
| Iron | Fe | 2+,3+ | I,III |
| Lead | Pb | 2+,4+ | II,IV |
| Tin | Sn | 2+,4+ | II,IV |
| Nickel | Ni | 2+,3+ | II,III |
| Gold | Au | 1+,3+ | I,III |

Write the name of the metal first, then add the name of the polyatomic ion. It'll usually end in "ate" Remember to add the relevant roman numeral after the metal

|  |  |  |  |
| --- | --- | --- | --- |
| Element | Symbol | Ionic Charges | Roman Numeral |
| Copper | Cu | 1+,2+ | I,II |
| Iron | Fe | 2+,3+ | I,III |
| Lead | Pb | 2+,4+ | II,IV |
| Tin | Sn | 2+,4+ | II,IV |
| Nickel | Ni | 2+,3+ | II,III |
| Gold | Au | 1+,3+ | I,III |

|  |  |
| --- | --- |
| Polyatomic Ion | Formula and Charge |
| hydroxide | OH­- |
| nitrate | NO3- |
| bicarbonate | HCO3- |
| chlorate | ClO3- |
| sulphate | SO4-2 |
| carbonate | CO3-2 |
| phosphate | PO4-3 |
| dichromate | Cr2O7-2 |
| ammonium | NH4+ |

Write the elements in the order they were initially written. To add prefixes, count the number of each element's atoms, and consult the chart. Remember to change the name of the last element, so it ends in "ide"

|  |  |
| --- | --- |
| # of atoms | Prefix |
| 1 | mono |
| 2 | di |
| 3 | tri |
| 4 | tetra |
| 5 | penta |
| 6 | hexa |
| 7 | hepta |
| 8 | octa |
| 9 | nona |
| 10 | deca |

"Special" molecular compounds

|  |  |
| --- | --- |
| Formula | Name |
| HNO3 | nitric acid |
| H2SO4 | sulphuric acid |
| HCl | hydrochloric acid |
| H2PO4 | phosphoric acid |
| NH3 | ammonia |
| O3 | ozone |
| H2O2 | hydrogen peroxide |
| H20 | water |
| CH4 | methane |

These compounds are used so frequently, that they have their own names (at least in SNC2D)

Write the name of the metal first, followed by the name of the non metal. The non metal's name should be changed to end in "ide"

Write the name of the metal first, then add the name of the polyatomic ion. It'll usually end in "ate"

|  |  |
| --- | --- |
| Polyatomic Ion | Formula and Charge |
| hydroxide | OH­- |
| nitrate | NO3- |
| bicarbonate | HCO3- |
| chlorate | ClO3- |
| sulphate | SO4-2 |
| carbonate | CO3-2 |
| phosphate | PO4-3 |
| dichromate | Cr2O7-2 |
| ammonium | NH4+ |

No

Is the compound ionic?

Yes

Yes

Yes

No

No

No

Does the

compound include any multivalent metals?

Does the

compound include any multivalent metals?

Does the

compound include any

polyatomic

Ions?

No

Is it a

"special" molecular compound?

End

Yes