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Full Length Article

[](http://crossmark.crossref.org/dialog/?doi=10.1016/j.ejbas.2018.10.001&domain=pdf)Structural, IR spectra NBO, TDDFT, AIM calculation, biological activity and docking property of [1,2,4]-triazolo[3,4-*b*][1,3,4] thiadiazole

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# a r t i c l e i n f o

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# a b s t r a c t

In this paper a complete quantum chemical study of [1,2,4]-triazolo [3,4-b][1,3,4] thiadiazole has been done with the combination of DFT/B3LYP method and 6-311G(d, p) basis set. The vibrational assignments are calculated with the help of PED. By using quantum theory of atoms in the molecule (QTAIM) we have calculated topological parameters at BCP point by which the nature of several intermolecular hydrogen bondings are studied. Electronic properties are calculated with the help of HOMO-LUMO plot. Chemical active sites of title molecule are described by FF, chemical hardness, chemical softness etc. UV spectra are calculated with the help of TDDFT using optimized parameters. NBO analysis gives idea about transfer of charge between bonding and antibonding electrons. Biological activity analysis suggests that the mole- cule can be used in anti- inflammatory drugs to explore new drugs. The molecule is docked with

MBNL1 receptor with the help of Swiss-Dock online server with Full fitness score of —541.58 a.u.

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1. Introduction

1,3-thiazole containing both Nitrogen and sulphur is a hetero- cyclic compound. The term ’thiazole’ refers to a large family of derivatives [[1]](#_bookmark18). Thiazole having molecular formula C3H3NS is a soft yellow liquid with pyridine-like scent [[2]](#_bookmark19) and used in a variety of specialized products, often by fusing with benzene derivatives. These are naturally occurring peptides, and utilized in the com- mercial development of peptidomimetics [[3,4]](#_bookmark21). Thiazole is hetero- cyclic series organic compound characterized by a ring structure composed of three carbon atoms, one nitrogen atom, and one sul- phur atom. This ring structure attributes many significant biologi- cally active natural qualities such as thiamine, penicillin, and in many synthetic drugs, dyes, and industrial chemicals [[5,6]](#_bookmark22). These are found in a number of various food products and flavors and are used in the manufacturing of prepackaged food products including noodles [[7]](#_bookmark23). Thiazoles help flavorists to create healthier crops by enhancing the flavor of meats and savory foods without

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the accumulation of excess salts or fats [[8]](#_bookmark24). 1,2,4-Triazole deriva- tives and their bonded heterocyclic correspondents are well known for their different biological activities and 1,2,4-triazole rings have been combined into ligands used in coordination compounds. The 1,2,4-triazole derivatives and their *N* bridged heterocyclic equivalents have been widely studied [[9–14]](#_bookmark25). Numerous 3,6-disubstituted [1,2,4] triazole [3,4-*b*] [1,3,4] thiadiazoles are testified to display significant antibacterial [[15–18]](#_bookmark26), pesticidal [[19]](#_bookmark28), anticancerous [[20]](#_bookmark28), anti-inflammatory, and anti-oxidant activities [[21]](#_bookmark28). Many studies are carried to synthesize organic com- pounds, inorganic metal complexes and are detected by using NMR FTIR spectra [[22–27]](#_bookmark28). Monirah A. Al-Alshaikh etal. synthesized the crystal structure of [1,2,4]-triazolo[3,4-b][1,3,4] thiadiazole molecule and found it as a potential bioactive agent [[22–28]](#_bookmark28). The present study on the [1,2,4]-triazolo[3,4-b][1,3,4] thiadiazole molecule has been carried out to support and as an extension of the work of Monirah A. Al-Alshaikh et al. [[29]](#_bookmark29). As mentioned in the work of Monirah A. Al-Alshaikh et al., due to bioactive potential of the molecule its biological properties are studied by using several parameters and also docking is made for designing new anti-inflammatory drugs. Along with this the geometrical parame- ters, NMA, Biological and Chemical activities, electronic transitions,

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thermodynamic properties are studied by the combination of DFT/ B3LYP method and 6-311G(d, p). A complete DFT analysis is per- formed on the title molecule ([Fig. 1](#_bookmark6)).

1. Computational details

All the calculations are done with G03 program package [[29]](#_bookmark29) using combination of B3LYP/DFT method and 6-311G (d, p) basis set. This basis set 6-311G (d, p), with ‘p’ polarization functions on hydrogen atoms and ‘d’ polarization functions on heavy atoms, is used for better description of polar bonds of molecule [[30,31]](#_bookmark30). Initial geometry is modeled with the help of Gauss View 3.0 pro- gram package [[32]](#_bookmark35). AIM calculation, performed by AIMALL pro- gram package [[33]](#_bookmark36), is based on quantum theory of atoms in a molecule. Natural bond analysis is carried with NBO 3.0 program package [[34]](#_bookmark37). Electronic transitions, vertical excitation energies, and oscillator strengths are computed with the TD-DFT method.

1. Result and discussion
   1. *Optimized parameters*

The molecule taken has two benzene rings, one thiazole and other phenyl rings. The molecule shows some deviation from pla- narity, with fluoro ring (C19-C14) forming dihedral angle of 2.65°

and the phenyl ring (C12-N17) 3.84°. The distance of the interac-

tion between benzene ring and thiazole ring C14-C19 is 2.33 Å.

AIM [[35]](#_bookmark39) calculation shows that there are two interactions between N18-H8 and N18-F29 in the molecule, which are shown in [Fig. 2](#_bookmark5). According to Koch and Popelier standard, an existence of hydrogen bond follows bond critical point (BCP) for the ‘proton (H).. .acceptor (A)’, [[36]](#_bookmark42) which involves electron density (q) in the

range of 0.002–0.040 a.u. and Laplacian (∇2 ) in the of range 0.024–

*q*

0.139 a.u. All calculated parameters at BCP, in between N18-H8,

and N18-F29, are listed in [Table 1](#_bookmark7). On the basis of topological parameters three types of H-bonds are characterized. According to Rendering to Rozas et al. [[37]](#_bookmark43), the characterization demands,

at BCP, that for strong and covalent nature of H-bond ∇2 < 0 and H < 0, for medium and partially covalent nature of H-bond ∇2 > 0 and H < 0, and for weak H-bond of electrostatic character ∇2 > 0 and H > 0. From [Table 1](#_bookmark7), the Laplacian of charge density are posi-

*q*

*q*

*q*

tive, however value of H < 0 signifies that the interaction is med- ium strong in nature. According to Espinosa et al. interaction energy calculated is Eint = ½ V at BCP [[38]](#_bookmark44). In this study calculated interaction energy of N18-H8, N18-F9 is 3.95 kcal/mol and 5.648 kcal/mol respectively.

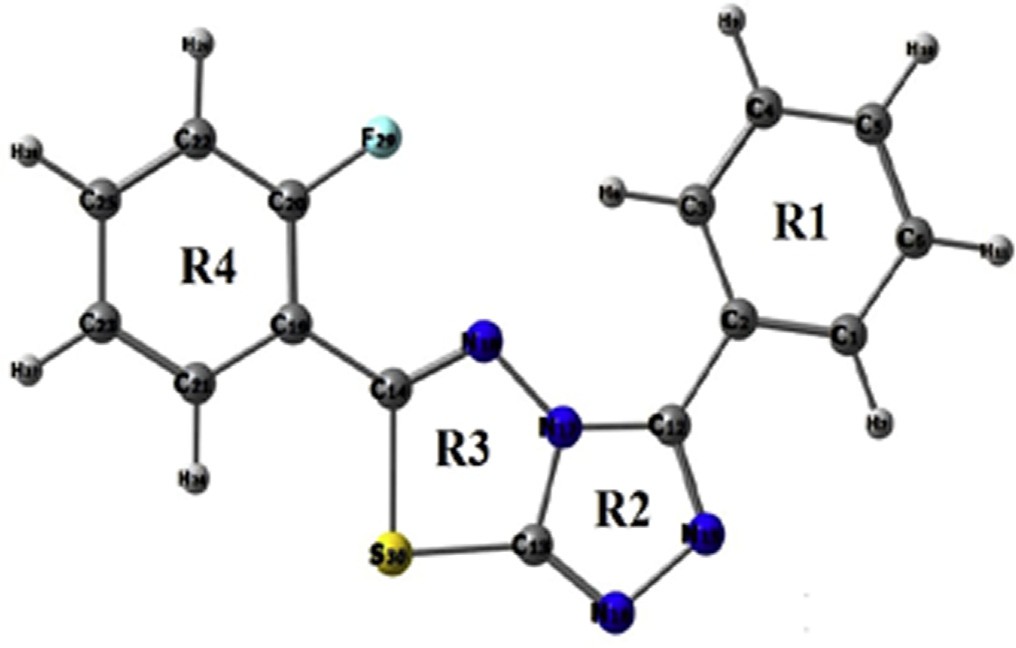


Fig 1. Molecular Picture of title molecule according to numbering.

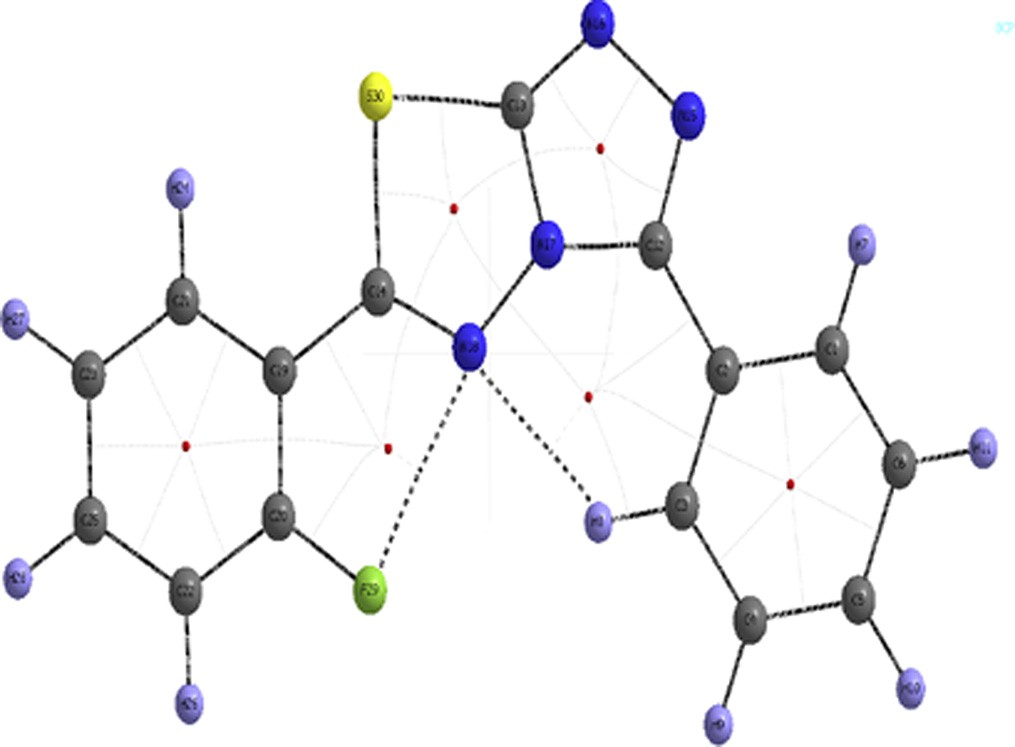


Fig 2. Molecular Graph of the studied compound at BCP.

* 1. *Electronic property and HOMO-LUMO surfaces*

The HOMO represents the ability to donate an electron by an atom while LUMO represents the ability to accept. The HOMO and LUMO energies are calculated by DFT/B3LYP method. The elec- tronic parameters, such as highest occupied molecular orbital (HOMO) energy (E-HOMO), lowest unoccupied molecular orbital (LUMO) energy (E-LUMO) and band gap energy (DE = FLUMO- FHOMO) are described theoretically [[39]](#_bookmark45). The energy (Egap) between HOMO and LUMO in the present study is found to be

3.71 eV. In this molecule LUMO is located over whole molecule however HOMO is located only on benzene ring. The transition from HOMO ? LUMO in the molecule indicates charge transfer from benzene ring to thiazole ring. The molecular electrostatic potential (MESP) is a map of potential i.e. electrostatic potential over constant electron density of all molecules. It displays molecu- lar size and shape as well as positive, negative, and neutral electro- static potential regions in terms of color grading scheme. In this color grading schemes, the blue color represents the most elec- tropositivity i.e. low electron region, whereas the red color corre- sponds to the most electronegative center or electron rich regions [[40,41]](#_bookmark48).The MESP plot of title molecule is display in [Fig. 3](#_bookmark8). From this fig. ring R2 is covered by red region so it is the most

electronegative region of title molecule. On the basis of energy of frontier orbitals, *e*HOMO, *e*LUMO [energy of highest occupied molecular orbital, Energy of lowest unoccupied molecular orbital respectively] the different global reactivity descriptors such as

electronegativity *v* (tendency to attract electrons), chemical poten-

tial *l* (energy that can be absorbed or released), global hardness *g* (hardness basically signifies the resistance against the deformation or polarization of the electron cloud of the atoms, ions or mole-

cules under small perturbation of chemical response), global elec- trophilicity index *x* (capacity of a species to accept arbitrary number of electrons) and global softness S (Softness signify how easily a molecule becomes polarized) are computed using equa-

tions as given below [[42]](#_bookmark50) ([Fig. 4](#_bookmark9)).

The energies of frontier molecular orbitals (*e*HOMO, *e*LUMO), energy band gap (*e*LUMO–*e*HOMO), electronegativity (*v*), chemical potential (*l*), global hardness (*g*), global softness (S), and global electrophilicity index (*x*) of the molecule have been listed in

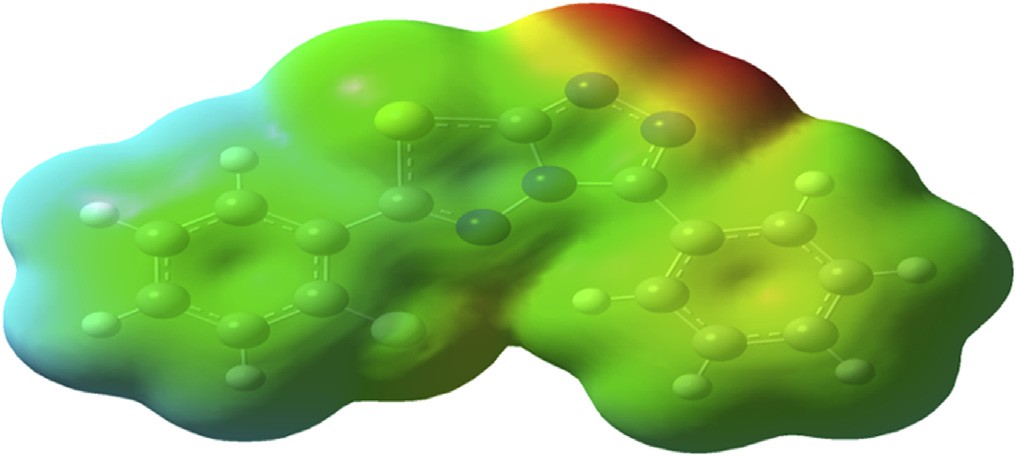
[Table 2](#_bookmark10). On the basis of *e*HOMO and *e*LUMO, these parameters are calculated by using equation given below [[43–45]](#_bookmark53)

## *v* = —1/2(*e*LUMO + *e*HOMO)

Table 1

Topological parameters for bonds of interacting atoms: electron density (*q*BCP), Laplacian of electron density (∇2BCP), kinetic electron energy density (G BCP), potential electron energy density (VBCP), total electron energy density (HBCP), estimated interaction energy (Eint) at bond critical point (BCP).

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| O–H | *q*BCP | ∇2 *q*BCP | *G*BCP | *V*BCP | *H*BCP | *E*int |
|  | (a.u.) | (a.u.) | (a.u.) | (a.u.) | (a.u.) | (kcal/mol) |
| N18-H8 | 0.0150 | 0.0630 | 0.0142 | —0.013 | —0.014 | 3.953 |
| N18-F29 | 0.0159 | 0.0897 | 0.0201 | —0.018 | —0.0186 | 5.648 |

* + 1. *Fukui function*

The Fukui function (FF) [[46,47]](#_bookmark31) of a molecule provides informa- tion about reactive sites and the method of understanding as well as categorizing chemical reactions. The atom with the highest FF value is highly reactive site as compared to the other atoms in the molecule. The use of the Fukui functions for the site selectivity in title molecule for nucleophilic and electrophilic attacks has been made with special emphasis on the dependence of the Fukui values with the basis of B3LYP/6-311G(d, p) level of theory. Using the Mullikan atomic charges of neutral, cation, and anion, state of

the molecule, the Fukui functions (*f* +, *f* —, *f* 0), local softness (*S*+,

*k k k k*

Fig. 3. MESP surfaces of thiazol.

*S*—, *S*0) [[48]](#_bookmark31) and local electrophilicity indices (*x*+, *x*—, *x*0 ) are cal-

*k k k k k*

culated using the following equations

## *f* + = [*qk* (N + 1)— *qk*(N)] for nucleophilic attack *f* — = [*qk* (N)— *qk* (N — 1)] for electrophilic attack *f* 0 = 1/2[*qk* (N + 1)+ *qk*(N — 1)] for radical attack.

*k*

*k*

*k*

where *qk* is electronic population of atom k in the neutral molecule with N electrons.

Local softness and electrophilicity indices are calculated using

*S* = S*f*, *So* = *Sf o*, *x* = *xf*, *xo* = *xf o*

*k k k k k k k k*

where +, —, and 0 signs show nucleophilics, electrophilics, and rad- ical attacks, respectively. The Fukui functions, for all the carbon

atoms in both benzene ring (R1 R4) and sulfar atom in thiazole ring (R3), calculated by using combination of DFT/B3LYP method and 6- 311G (d, p) basis set are listed in [Table 3](#_bookmark10). From [Table 3](#_bookmark10), it can be seen that the carbon of both benzene ring, 14C and 31C are more reactive sites for nucleophilic and 26C as an electrophilic reactive attacks. However 30S atom of the thiazole ring shows more favor- able site for both electrophilic and nucleophilic substitution than ring carbon atoms. Furthermore, 30S site is more reactive to the

electrophilic than nucleophilic substitution because *f* — > *f* +.

*k k*

Fig. 4. HOMO, LUMO surfaces of thiazol.

*l* = —*v* = 1/2(*e*LUMO + *e*HOMO)

## *g* = 1/2(*e*LUMO — *e*HOMO)

S = 1/2*g x* = *l*2 /2*g*

Time dependent density functional theory (TD–DFT) method is important tool for studying the nature of the transitions of UV–vis spectrum of the compound. Optimized geometry is obtained for TD-DFT calculation by using same level theory. UV spectra of the molecule, calculated by combination of DFT/B3LYP method and 6-311G (d, p) basis set is shown in [Fig. 6](#_bookmark12). In this graph, two promi- nent peaks observed at 4.02 eV (294 nm), 2.03 eV (610 nm) origi- nate mainly due to H ? L (37%), H-4 ? L (43%) transitions respectively. These transitions are shown in [Fig. 5](#_bookmark11) and listed in [Table 4](#_bookmark13). On the basis of the calculated molecular orbital coefficients

analyses electronic transition are assigned to n ? *p*\* and p ? p\*

respectively.

* 1. *Non linear optical parameter*

The Gaussian 03 program is used to calculate the dipole moment (*l*) and polarizability (*a*) of the molecules, based on the finite field approach. The first hyperpolarizability (*b*) and polariz- ibility of the title molecule are calculated by using same level the-

Table 2

Electronic parameters of title compound calculated by DFT/B3LYP method and 6-311G(d, p).

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| HOMO energy | LUMO energy | Electronegativity (*v*) | Chemical potential (*l*) | Global hardness (*g*) | Global softness (S) | Global electrophilicity index (*x*) |
| —6.255 | —2.545 | 4.400 | —4.400 | 1.855 | 0.270 | 5.218 |

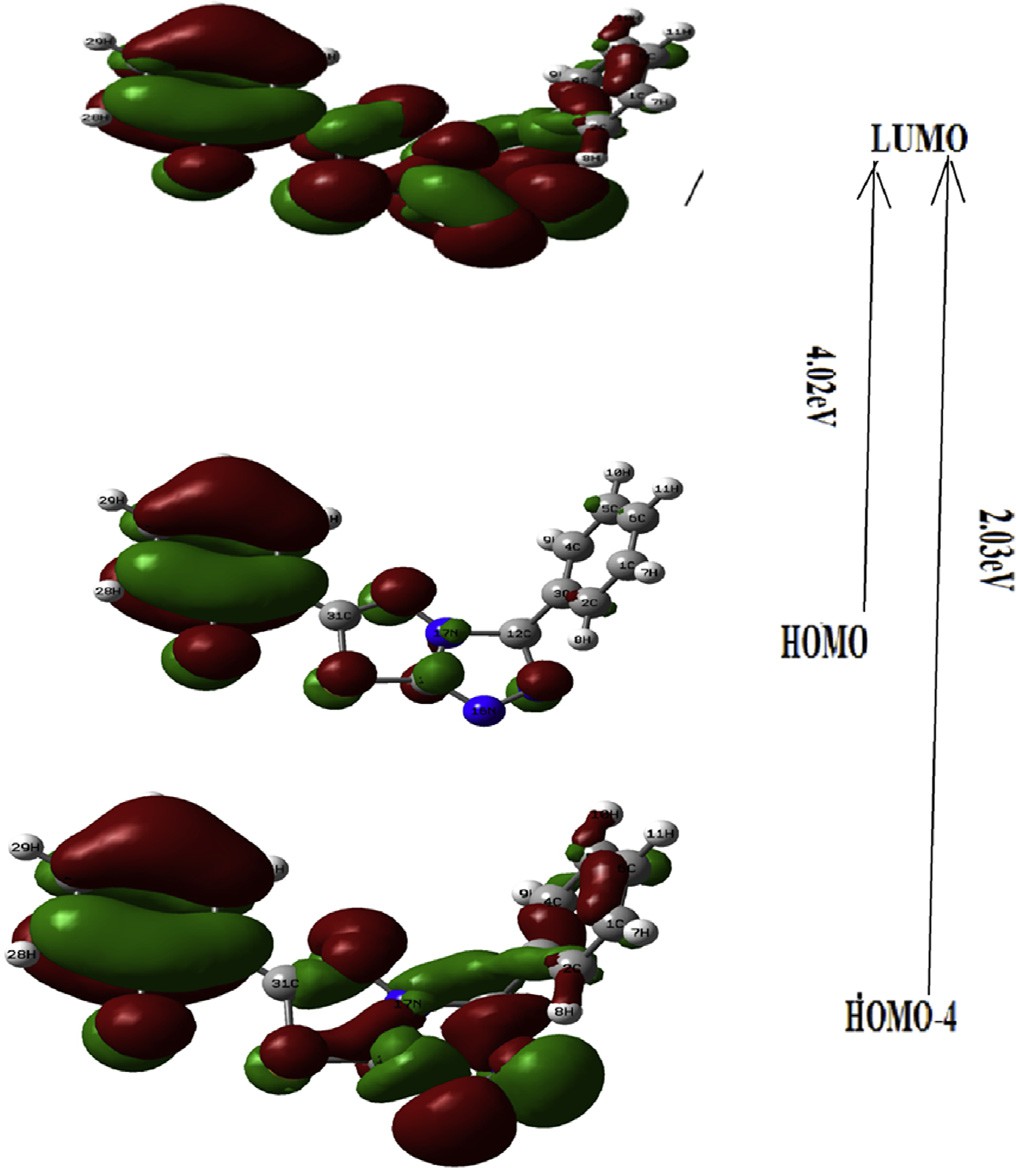
Table 3

Fukui values on the basis theory of the Mulliken atomic charges of neutral, cation, and anion, the Fukui functions (*f* +, *f* —, *f* 0).

*k k k*

*k k*

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| S.N | Neutral | Positive | Negative | *f* + | *f* — | *f* 0  *k* |
| 1C | —0.133 | —0.128 | —0.136 | 0.005 | 0.003 | —0.132 |
| 2C | —0.171 | —0.173 | —0.163 | —0.002 | —0.008 | —0.168 |
| 3C | 0.188 | 0.190 | 0.191 | 0.002 | —0.003 | 0.191 |
| 4C | —0.171 | —0.169 | —0.169 | 0.002 | —0.002 | —0.169 |
| 5C | —0.133 | —0.129 | —0.135 | 0.004 | 0.002 | —0.132 |
| 6C | —0.123 | —0.113 | —0.129 | 0.010 | 0.006 | —0.121 |
| 12C | 0.041 | —0.011 | 0.077 | —0.052 | —0.036 | 0.033 |
| 14C | 0.313 | 0.344 | 0.239 | 0.031 | 0.074 | 0.292 |
| 20C | 0.069 | 0.063 | 0.070 | —0.006 | —0.001 | 0.066 |
| 21C | —0.160 | —0.147 | —0.183 | 0.013 | 0.023 | —0.165 |
| 22C | 0.313 | 0.332 | 0.285 | 0.019 | 0.028 | 0.309 |
| 23C | —0.132 | —0.126 | —0.134 | 0.006 | 0.002 | —0.130 |
| 25C | —0.190 | —0.181 | —0.199 | 0.009 | 0.009 | —0.190 |
| 26C | —0.124 | —0.105 | —0.154 | 0.019 | 0.030 | —0.130 |
| 31C | 0.039 | 0.065 | 0.002 | 0.026 | 0.037 | 0.034 |
| 30S | 0.301 | 0.484 | 0.118 | 0.183 | 0.183 | 0.301 |

ory and listed in [Table 5](#_bookmark14). In the presence of applied electric field, energy is function of applied electric field. Hyperpolarizability and Polarizability are described as the response of a system due to application of the electric field [[49]](#_bookmark31). The Hyperpolarizability and Polarizability are determined by nonlinear optical properties (N L O) of the system as well as strength of molecular interaction [[50,51]](#_bookmark31). First order Hyperpolarizability of the system is defined

by 3 × 3 × 3 matrix of is a third rank tensor. The third rank matrix

which contains 27 components can be reduced to 10 components

by using Kleinman symmetry [[52]](#_bookmark31). All components of *b* are coeffi- cients in the Taylor series expansion of the energy. When weak and homogeneous external electric field is applied, this expansion becomes

*E* = *E*0—*laFa* — 1/2*aabFaFb* — 1/6*babcFaFbFc*

where *E*0 is the energy of the unperturbed molecules, *Fa* the field at

the origin *la*, *aab* , and *bab* are the components of dipole moment, polarizability and the first hyperpolarizabilities respectively.

The total dipole moment and the mean polarizability and hyperpolarizability using x,y,z components are defined as

*l* = (*l*2 + *l*2 + *l*2)1/2

*x y z*

where *lx* , *ly* , *lz* are X,Y,Z components of dipole moment

⟨*a*⟩= 1/3[*axx* + *ayy* + *azz*]

where *axx* , *ayy* , *azz* are the components of polarizibility in XX, YY, ZZ

planes

Fig 5. Electronic Transition of TD-DFT of title molecule by DFT/B3LYP method.

*b* = (*b*2 + *b*2 + *b*2 1/2

*Total*

*x y z* )

2 2 2 1/2

= [(*b* + *b* + *b* ) + (*b* + *b* + *b* ) + (*b* + *b* + *b* ) ]

*xxx xyy xzz yyy yxx yzz zzz zxx zyy*

and in bc plane and (1 a.u. = 3.6998 × 10—30 e.s.u.). where *babc* are components of hyperpolarizibility along a direction

The calculated value of dipole moment is 6.29 D. In title mole-

cule *axx* and *ayy* have greater contribution of polarizibility which indicates that the molecule is more polarized along Y and X direc- tions. The calculated value of hyperpolarizibily of given molecule is

nearly same as, 2-{[5-(adamantan-1-yl)-4-methyl-4H-1,2,4-tria zol-3-yl]sulfanyl}-N,N-dimethylethanamine [[53]](#_bookmark31), and nearly one fifth of Thiazole Azo Dyes.[[54]](#_bookmark31) Urea is benchmark compound for NLO properties, therefore we compare our results with urea. The calculated value of hyperpolarizibily of the compound is nearly nine times greater than hyperpolarizibily of urea so it is a good claimant as a NLO material.

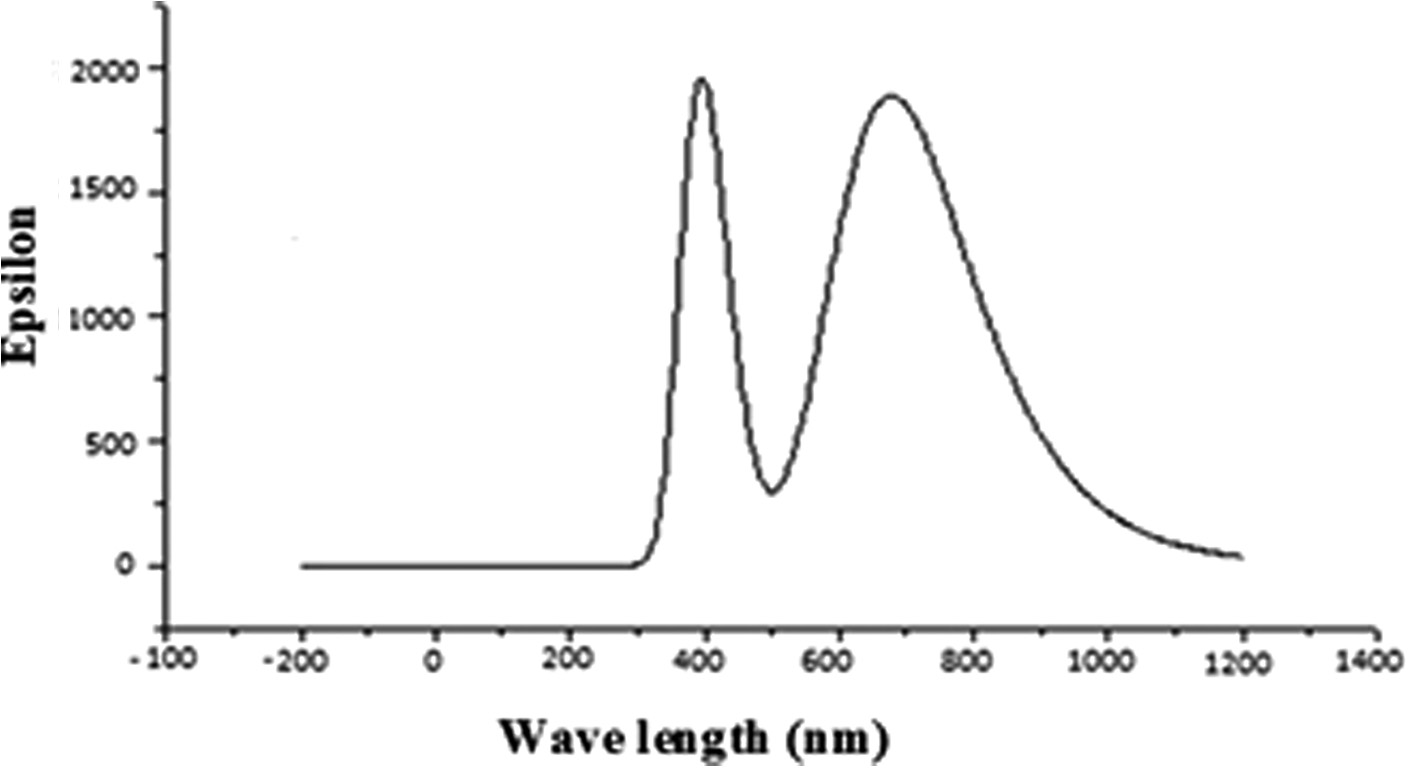


Fig 6. UV spectra of title molecule calculated by combination of DFT/B3LYP method and 6-311G(d,p) basis set.

Table 4

The observed UV–vis spectra as calculated by TD-DFT method at B3LYP/6-311 G (d, p) level.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Excitation energy (eV) | Wavelength (nm) |  | Oscillator strength | Orbital transition |
|  | Calculated | Assignment |  |  |
| 4.02 | 294 | n ? *p*\* | 0.023 | HOMO ? LUMO (43%) |
| 2.03 | 610 | *p* ? *p*\* | 0.130 | HOMO-4 ? LUMO (37%) |

Table 5

Polarizability and Hyper Polarizability of title compound calculated by same level theory.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| S.N. | Parameter | Polarizability | S.N. Parameter | Hyper Polarizability |
| 1. | *axx* | —113.889 | 1. *bxxx* | —130.786 |
| 2. | *ayy* | —129.426 | 2. *byyy* | 58.531 |
| 3. | *azz* | 0.393 | 3. *bzzz* | 4.053 |
| 4. | *axy* | 6.624 | 4. *bxyy* | —42.254 |
| 5. | *axz* | 2.339 | 5. *bxxy* | 77.013 |
| 6. | *ayz* | —2.853 | 6. *bxxz* | 35.379 |
| 7 | *a* | —9.678 | 7. *bxzz*  8. *byzz* | 14.626  —1.132 |
|  |  |  | 9. *byyz* | —1.432 |
|  |  |  | 10. *bxyz* | —1.856 |
|  |  |  | 11. *b* | 211.201 |

* 1. *Vibrational analysis*

The molecule under study contains 30 atoms, so there are 3N-6 modes of vibrations. All mode of vibration are studied with the help of combination of Gauss View and VEDA 4.0 program package [[55]](#_bookmark31). In our calculations we ignore electron- electron correlation as well as an-harmonicity [[56,57]](#_bookmark31). These approximations cause calcu- lated frequencies to be in higher region than experimental results, therefore we scale calculated frequency by 0.9688. [[58]](#_bookmark32) Some of selected modes of vibrations are listed in [Table 6](#_bookmark15) and calculated IR spectra of the molecule is also plotted in [Fig. 7](#_bookmark16), using same level theory. Some important modes of vibrations are discussed below.

* + 1. *C-H stretching*

stretching vibrations in the region 2800–3100 cm—1 [[58]](#_bookmark32). In the The hetero aromatic structure shows the presence of C–H is observed at 3066–3108 cm—1. Two sharp back to back peaks present study the C–H stretching vibration of the title compound

observed at 3077 cm—1, 3076 cm—1 due to *t*(C-H)R1 and mixing of

*t*(C-H)R1,R4 with PED 24%, 43% respectively, are well matched with

nyl)benzo[d]thiazole at 3076 cm—1 [[59]](#_bookmark33). An intense mixing of modes are observed at 1625 cm—1 due to *b* (CH) R1 with PED experimental as well as calculated IR spectra of 2-(4-methoxyphe

peaks are observed at 1611 cm—1 and 1605 cm—1 due to *b* (CH)R1, 16%. In mid region of spectra very intense and sharp polarized *b*(CH)R4 respectively with PED 47% and 76%. In lower region some

out of plane bending and twisting mode of vibrations occur.

* + 1. *C-N vibrations*

The determination of C-N vibrations is a difficult task, because of the mixing of vibration in this region. Silverstein [[58]](#_bookmark32) has first

1266 cm—1. The C-N stretching vibration (15N-12C) and (17N- 12C) in the molecule observed at 1266 cm—1 and 1263 cm—1 with time assigned C-N stretching absorption in the range of 1382–

(1267 cm—1) with C-N stretching mode and lies in the same region PED contribution 15% and 46% respectively, are well matched as 2-(4-methoxyphenyl)benzo[d]thiazole [[59]](#_bookmark33). In lower rage of fre-

17N) and torsion (14C-18N-17N-12C) are obtained at 1088 cm—1 and 654 cm—1 respectively with PED contribution 20% and 79%. quency some other in plane bending vibration modes (14C-18N-

Table 6

Theoretical (selected) vibrational wave numbers of title molecule using B3LYP/6– 311G (d, p) and their assignment.

Cal. Freq. Sc. IR Int. Mode of vibration freq

ring experiences torsion *s* (14C-17N-18N-30S) at 936 cm—1 with PED 21%, however opposite half segment of thiazole ring experi-

520 cm—1 with PED 27%. In fuoro ring a medium intense polarized ences very intense peak of torsion s (29F-19C-22C-20C) at due to *t* (29F-20C), is obtained at 1207 cm—1 with PED 15%. Some mode of vibration with polarization vector along the plane of ring, other modes of vibrations are reported in table however those

mode of vibration are not reported for which PED < 15%.

* 1. *Donor acceptor interaction*
     1. *NBO analysis*

NBO gives the most suitable ‘natural Lewis structure’ because all orbitals include the highest possible percentage of the electron density (ED). NBOs deliver an accurate method for studying intra and intermolecular interaction as well as charge transfer or con- jugative interaction in different molecular systems [[60]](#_bookmark34). The value of *E*(2) shows the strength of interaction in between electron donors and electron acceptors, i.e., As the value of *E*(2) increases the strength of interaction increases and vice-versa. For each donor NBO(*i*) and acceptor NBO(*j*), the strength of delocalization interac- tion (or stabilization energy) *E*(2) associated with electron delocal- ization between donor and acceptor is estimated by the second order energy lowering as [[61,62]](#_bookmark38)

*F*2

*E*(2) = D*Eij* = *qi*

|  |  |  |  |
| --- | --- | --- | --- |
| 3176 | 3077 | 22.26 | R1[*t* (1C-7H) (24%)] |
| 3175 | 3076 | 4.35 | R1[*t* (6C-11H) (24%)] |
| 3172 | 3073 | 33.69 | R4[*t* (26C-29H)] (27%) + R1[*t* (1C-7H)] (16%)] |
| 3168 | 3069 | 4.36 | R1[*t* (4C-9H) (47%) + *t* (5C-10H) (26%)] |
| 3154 | 3056 | 3.89 | R4[*t* (25C-28H) (50%)] |
| 3152 | 3054 | 3.89 | R1[*t* (1C-7H) (49%)] |
| 3148 | 3050 | 9.39 | R4[*t* (21C-24H) (37%)] + R1[*t* (6C-11H) (26%)] |
| 1677 | 1625 | 540.56 | R1[*b* (2C-3C) (11%)] + R1[*b* (1C-2C) (14%)] |
| 1663 | 1611 | 581.20 | R1[*t* (2C-3C) (47%)] |
| 1657 | 1605 | 34.93 | R4[ *b* (24H-21C)(42%)] + R4[*b* (22H-23C) (34%)] |
| 1637 | 1586 | 40.23 | R1[*t* (1C-2C)[31%)] + R2[*t* (15 N-12C) (11%)] |
| 1622 | 1571 | 52.34 | R1[*t* (1C–2C) (29%)] + R1[*b* (6C-11H)(34%)] |
| 1577 | 1528 | 14.89 | R1[*t* (5C-4C) (14%)] + R1[*b* (5C-4C-6C) (15%)] |
| 1540 | 1492 | 19.68 | R1[*b* (1C-2C-3C) (14%)] |
| 1520 | 1473 | 14.30 | R1[*b* (8H-2C) (17%)] + R4[*b* (21C-23C-26C) (21%)] |
| 1493 | 1446 | 0.12 | R1[*b* (10H-5C-6C) (11%)] + R1[*b* (8H-2C-1C) (13%)] |
| 1477 | 1431 | 150.32 | R1[*t* (1C-2C) (27%)] |
| 1411 | 1367 | 29.25 | R3[*t* (14C-30S) (17%)] |
| 1408 | 1364 | 34.13 | R1[*t* (2C-3C) (15%)] + R1[*b* (8H-2C-1C) (21%)] |
| 1383 | 1340 | 2.43 | R1[*c* (3C-2C-1C) (21%)] |
| 1341 | 1299 | 9.29 | R2[*t* (14C-18N) (19%)] + R2[*t* (15 N-12C) (13%)] |
| 1327 | 1286 | 3.30 | R2[*t* (12C-17N) (27%)] + R1[ *b*(3C-6C-5C) (15%)] |
| 1326 | 1285 | 3.24 | R2[*t*(15N-12C) (39%)] |
| 1287 | 1247 | 10.02 | R2[ *b*(17N-12C-15N) (14%)] + R3[*b*(18N-14C-30S) |
|  |  |  | (23%)] |
| 1254 | 1215 | 150.36 | R1[ *t*(2C-3C) (18%)] + R2[*b*(12C-15N-16N) (13%)] |
| 1218 | 1180 | 15.79 | R2[ *t*(17N-14C) (21%)] + R3[*c*(14C-18N-17N) (19%)] |
| 1187 | 1150 | 64.61 | R3[*t*(14C-30S) (17%)] + R1[*c*(11H-6C-5C) (23%)] |
| 1176 | 1139 | 0.26 | R2[*b*(17N-12C-15N) (23%)] + R3[ *b*(17N-14C-18N) |
|  |  |  | (21%)] |
| 1150 | 1114 | 15.86 | R4[ *c*(26C-23C-21C) (15%)] + R1[ *c* (1C-2C-3C)  (21%)] |
| 1119 | 1084 | 5.45 | R2[*t*(15N-12C) (13%)] + R4[*b*(22C-29F) (42%)] |
| 1107 | 1072 | 0.73 | R3[ *b*(18N-14C-30S) (46%)] + R1[ *b*(3C-2C-1C) (15%)] |
| 1002 | 971 | 0.45 | R2[*b*(17N-12C-15N) (12%)] + R1[*b*(1C-2C-3C) (11%)] |
| 976 | 946 | 5.62 | R3[*t*(30C-18N) (15%)] |
| 966 | 936 | 6.31 | R4[*s*(20C-21C-22C-23C) (24%)] + R3[*s*(14C-17N-  18N-30S) (21%)] |
| 925 | 896 | 30.44 | R2[*t* (15N-12C) (22%)] |
| 865 | 838 | 0.71 | R4[*s*(22C-23C-25C-26C) (11%)]+R2[*s*(17N-12C-  15N-16N) (10%)] |
| 850 | 822 | 1.05 | R2[*b*(14C-17N-18N) (13%)] |
| 776 | 752 | 10.88 | R1[ *s*(3C-6C-1C-2C) (11%)] + R3[*s*(14C-17N-18N-  30S) (15%)] |
| 660 | 639 | 5.24 | R1[*c*(6C-5C-4C) (11%)] + R2[*c*(12C-15N-16N) (17%)] |
| 655 | 635 | 6.47 | R1[*s*(3C-6C-2C-1C) (24%)] + R3[*s*(14C-18N-17N-  13C) (18%)] |
| 606 | 587 | 0.98 | R4[*c*(19C-20C-22C) (17%)] + R1[*c*(1C-2C-3C) (13%)] |
| 531 | 514 | 279.57 | R3[*s*(13C-17N-18N-30S) (30%)] + R1[*s*(6C-5C-4C-  3C) (42%)] |

*ij*

*ei* — *ej*

*Abbreviations*: *b* = in plane bending, *c* = out of bending, *x* = wagging, *s* = twisting, S = scissoring, R = rocking, *t* = stretching. Note: The PED distribution less than 12% are neglected for the sake of calculation.

* + 1. *C-C vibrations*

Normally the (C-C) stretching vibrations are expected within the middle region of spectra. The C–C stretching vibrations of the aromatic compounds, calculated by B3LYP/6-311G (d, p) method,

are at 1600 cm—1, 1586 cm—1, 1568 cm—1, 1500 cm—1 and 1293 cm—1 in thiazole ring however in benzene ring C–C aromatic stretching mode of vibrations lies in between 1645 and 847 cm—1. The other modes of vibration i.e. in-plane and out-of-plane bend-

ing are also observed at lower frequencies.

* + 1. *Other modes of vibration*

Some other modes of vibrations are also reported at lower range of frequencies. In thiazole ring intense plane polarized mode of vibration due to *t* (29F-20C) with PED 19% is obtained at

1207 cm—1. In this study below 1000 cm—1 half segment of thiazole

where *qi* is the population of donor orbital or donor orbital occu-

pancy;*ei* , *ej* are orbital energies (diagonal elements) of donor and acceptor NBO orbitals respectively; *Fij* is the off–diagonal Fock or Kohn–Sham matrix element between *i* and *j* NBO orbitals. The second-order perturbation analysis of Fock matrix in NBO basis are listed in [Table 7](#_bookmark17). From this table we see that strong intra-

molecular hyper conjugative interactions are formed by orbital overlap Lp(1)N15/N17,Lp(2)S30,*r*(C-C),*r*(C-H)p(C-C)*r*\*(C-C),

*r*\*(C-H), *p*\*(C-C) which results in intra-molecular charge transfer causing stabilization of the system. From [Table 7](#_bookmark17), interactions between C1-C2 to C3-C4, C5-C6 of *p*(C1-C2) ? *p*\*(C3-C4) and *p*(C1-C2) ? *p*\*(C3-C4) with electron density 0.9897 e stabilize by

10.28 kcal/mol, 10.49 kcal/mol respectively. The interaction between C3-C4 to C1-C2, C4-C5, C12-C17 of *p*(C3-C4) ? *p*\*(C1-C2)

*p*(C3-C4 ? *p*\*(C5-C4), *p*(C3-C4) ? *r*\*(C12-C17) with electron den- sity 0.833e stabilizes by 10.04 kcal/mol, 9.72 kcal/mol and

3.11 kcal/mole respectively. Similarly interaction between C5-C6 to C3-C4, C1-C2, of *p*(C5-C6) ? *p*\*(C3-C4) *p*(C5-C6) ? *p*\*(C1-C2)

with electron density 0.829 e stabilizes by 9.6 kcal/mol, 10.54 kcal/mol respectively. A strong interaction observed in between lone pair LP(1)N17 ? *p*\*(C14-N18), *p*\*(N18-C31) with energy 21.09 kcal/mol

and 20.25 kcal/mol respectively are comparable with interaction has been observed due to the electron density transfer from the lone pair (LP1) of nitrogen atom to antibonding orbitals *p*\*(C–N)

and *p*\*(N–C) with stabilization energies 41.07 and 46.78 kcal/mol,

respectively in similar molecular system 2-{[5-(adamantan-1-yl)-4

-methyl-4H-1,2,4-triazol-3-yl]sulfanyl}-N,N-dimethylethanamine [[53]](#_bookmark31). Since *p* orbitals have lower occupancies than r, correspond- ingly showing more electron- donating ability in comparison to *r*

orbital. This also shows that [LP(1) (S30) ? *p*\*(C14-C16)] is the most intensive interaction between the acceptor and donor which results in the molecular stability. A very strong interaction by *p*\* to *p*\* orbital overlap of *p*\*(C22-C25) ? *p*\*(C23-C26) stabilizes by

153.69 kcal/mol. In this system stabilization occurs by intermolecu- lar interactions of orbital overlap among *p* and *p*\* orbitals, which results in intermolecular charge transfer (ICT). Since ICT makes molecule more polarized, it must be responsible for the NLO

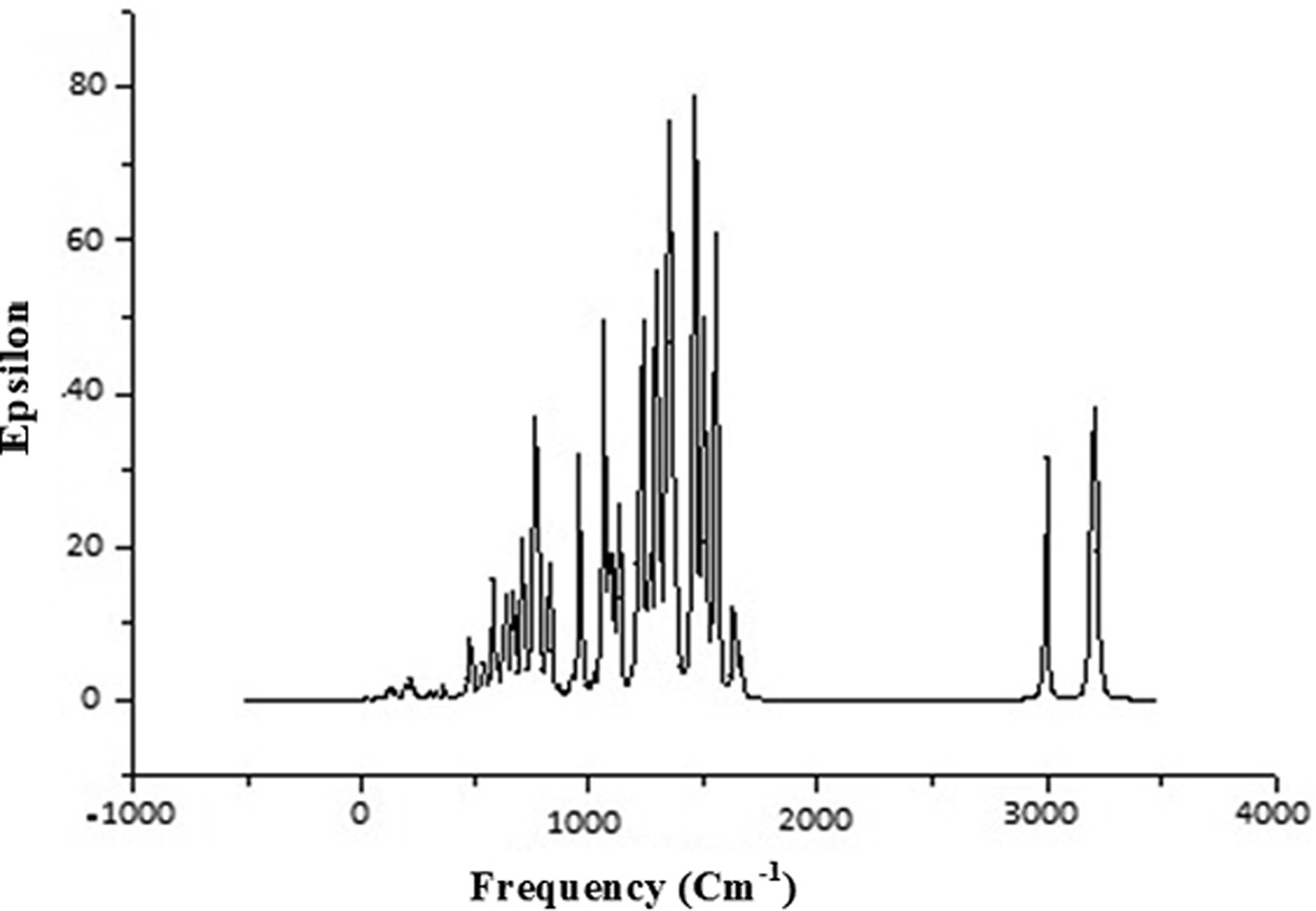


Fig 7. Calculated IR spectra of title molecule by using DFT/B3LYP method.

Table 7

Second–order perturbation theory analysis of the Fock matrix, in the NBO basis for intermolecular interactions within title molecule: Stabilization energy of interactions (*E*(2)), Energy difference between donor (*i*) and acceptor (*j*) NBO orbitals (*Ej*–*Ei*), Fock matrix element between *i* and *j* NBO orbitals (*Fij*).

|  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- |
| Donor NBO | Occupancy | Acceptor NBO | Occupancy *E*(2) (*Ej*–*Ei*) *F*(*i,j*) | | | |
| (*i*) | (*i*) | (*j*) | (*j*) | (kcal/mol | a.u | a.u. |
| *p* (C1-C2) | 0.9897 | *p*\*(C3-C4) | 0.1792 | 10.28 | 0.28 | 0.07 |
| *p*(C1-C2) | 0.9897 | *p*\*(C5-C6) | 0.1632 | 10.49 | 0.28 | 0.07 |
| *r*(C1-C6) | 0.9906 | *r*\*(C1-C2) | 0.0070 | 1.26 | 1.28 | 0.05 |
| *r*(C1-H7) | 0.9914 | *r*\*(C2-C3) | 0.0118 | 1.88 | 1.10 | 0.06 |
| *r*(C1-H7) | 0.99140 | *r*\*(C5-C6) | 0.0078 | 1.78 | 1.10 | 0.06 |
| *p*(C3-C4) | 0.8334 | *p*\*(C1-C2) | 0.1552 | 10.04 | 0.29 | 0.07 |
| p(C3-C4) | 0.8334 | p\*(C5-C6) | 0.1632 | 9.72 | 0.28 | 0.07 |
| p(C3-C4) | 0.8334 | r\*(C12-N17) | 0.0308 | 3.11 | 0.55 | 0.06 |
| p(C5-C6) | 0.8296 | p\*(C1-C2) | 0.1552 | 9.60 | 0.28 | 0.07 |
| p(C5-C6) | 0.8296 | p\*(C3-C4) | 0.1792 | 10.54 | 0.28 | 0.07 |
| Lp(2) N15 | 0.9998 | p\*(C14-S30) | 0.0226 | 22.72 | 0.21 | 0.09 |
| Lp(1) N17 | 0.9996 | p\*(C14-N18) | 0.2496 | 21.09 | 0.27 | 0.10 |
| Lp(1) N17 | 0.9996 | r\*(N18-C14) | 0.0337 | 20.25 | 0.26 | 0.09 |
| Lp(2)S30 | 0.9997 | p\*(C14-N18) | 0.0138 | 11.85 | 0.25 | 0.07 |
| Lp(2)S30 | 0.9997 | p\*(N18-C30) | 0.2145 | 9.85 | 0.25 | 0.06 |
| Lp(3)F29 | 0.9999 | p\*(C22-C25) | 0.0135 | 9.34 | 0.43 | 0.09 |
| p\*(N17-C13) | 0.1732 | p\*(C12-C2) | 0.0245 | 28.45 | 0.04 | 0.07 |
| p\*(C22-C25) | 0.1840 | p\*(C23-C25) | 0.0165 | 153.69 | 0.01 | 0.08 |

properties of molecule. Consequently the compound may be used for NLO materials in future.

* 1. *Biological properties*

Before discussing the biological activity of the molecule aque- ous solubility (log *S*) and Lipophilicity (log *P*) are calculated. For evaluating log *P* and log *S,* ALOGPS 2.1 program [[63]](#_bookmark40) is used. This program developed by Tetko et al. [[64–66]](#_bookmark41) is based on electrotopo- logical state indices and associative neural network modeling. The Structure-property relationship studies are justified by these two parameters. The transport property of drugs and their interaction with receptors are closely related by Log *P* however bioavailability is closely related by log *S*. The calculated value of log *P* is 2.34 which suggests that molecule is able to diffuse across the cell membranes, so it can be utilized in pharmacological applications. In general 85% of drugs have log *S* values in between the range of

—1 to —5.36. The calculated value of log *S* (—3.30) further approves the permeability of molecules through cell membranes. Some bio-

logical activities of title molecule are calculated with the help of PASS software. PASS calculates 900 pharmacological properties, molecular mechanisms of action, mutagenicity, carcinogenicity, teratogenicity, and embryotoxicity. PASS predicts these properties on the basis of structure activity relationships for the training set, considering more than 46 000 drugs, drug candidates, and prime compounds whose biological activity are resolute experimentally. The calculated results by PASS software have average accuracy of about 85% [[67]](#_bookmark46). In this study we only listed those biological activ- ities, in [Table 8](#_bookmark20), for which Pa > 70%. The molecule shows high value of biological activity for Chloride peroxidase inhibitor. The Chloride peroxidase belongs to the family of enzymes, which catalyzes the chlorination of organic compounds. It employs one cofactor, which may be either heme or vanadium [[68]](#_bookmark47). The molecule inhibits chlo- rination of organic compounds in blood stream which prevents

Table 8

Some biological activities calculated by PASS with Pa > 70%.

|  |  |  |
| --- | --- | --- |
| S.N. Biological activity | Pa | Pi |
| 1. Chloride peroxidase inhibitor | 0.756 | 0.005 |
| 2. Albendazolemonooxygenase inhibitor | 0.688 | 0.004 |
| 3. Corticosteroid side-chain-isomerase inhibitor | 0.682 | 0.010 |
| 4. Phobic disorders treatment | 0.686 | 0.082 |
| 5. Glucocorticoids | 0.646 | 0.050 |

bacterial disease. It is used for Phobic disorders treatment also. Phobias are the most common type of anxiety disorder, in which a person experiences an extreme or irrational fear of a place, object, animal, or situation. Corticosteroid side-chain isomers exist in all organs but show highest activity in liver and kidneys. The molecule inhibits Corticosteroid side-chain isomers in liver and kidneys and protect these organ from multiple myeloma. Gluco- corticoids (0.646) are part of the response mechanism in the im- mune system which decreases certain features of immune function, such as drop of inflammation [[69]](#_bookmark49). So they are used as medicines to treat diseases caused by an intense resistant sys- tem. They also inhibit growth in cancer cells, therefore high doses of Glucocorticoids are used to treat cancer. This contains inhibitory effects on lymphocyte production, as in the handling of lym- phomas and the moderation of side effects of anticancer drugs. Designing new anti-inflammatory in Myotonic dystrophy diseases, agents require the identification of targets, which when inhibited can kill the affected cells. We have done molecular docking studies by using Swiss-Dock web server [[70]](#_bookmark51). MBNL1 receptor is identified as an efficient target, based on the prediction by Swiss-Dock web server. Muscleblind-like (Drosophila), also known as MBNL1, is a protein which is encoded in humans by the MBNL1 gene [[71–](#_bookmark52) [73]](#_bookmark52). MBNL1 is stated in the early heart, and its levels rise across fatal and postpartum development. The docking score is expected to have minimum interaction energy value (e-value). As more neg- ative the e-value is the docking becomes more efficient. Docking process does not cover on a specific region but over entire protein. The docking picture of the molecule, obtained from the UCSF chi- mera software, is shown in [Fig. 8](#_bookmark27). The Full Fitness score for the title

molecule is —541.58 a.u. and binding affinity (DG in kcal/mol)

—5.46 a.u., which suggest that it has good binding affinity. The results obtained from docking studies suggest that the compound

can be used as an anti-inflammatory agent in Myotonic dystrophy.

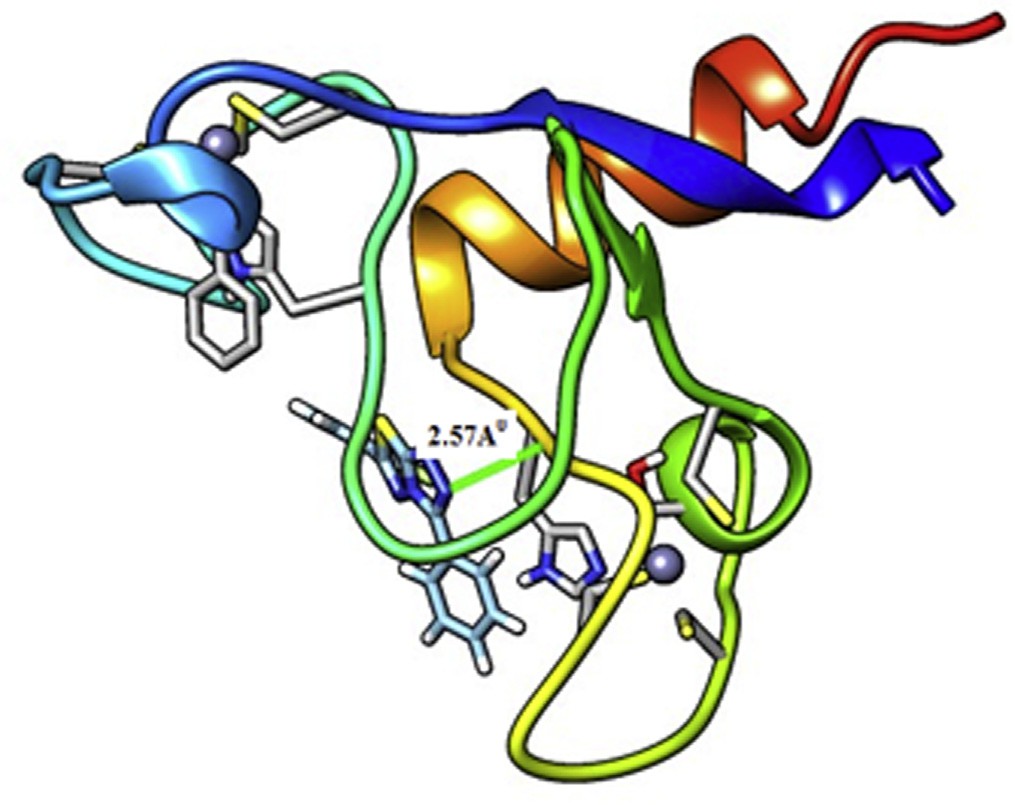


Fig 8. Molecular Docking of MBNL1 protein with title molecule.

1. Conclusion

Equilibrium geometries, harmonic frequencies, electronic prop- erties, NBO, NLO analysis, TDDFT and biological activity of title molecule are determined and analyzed using combination of DFT/B3LYP level of theory using 6-311G (d, p) basis set. The theo- retical calculation of UV–vis spectra shows a medium intense peak (at 610 nm) corresponding to H-4 ? L. The NBO analysis shows that strongest interaction with energy 153.69 kcal/mol corre-

sponding to *p*\*(C22-C25) ? *p*\*(C23-C26) stabilizes flouro ring.

The chemical reactivity of the molecule is explained by energy gap (3.71 eV) and plot of HOMO-LUMO. FF shows that 14C and 31C are better nucleophilic centers however 30S shows better elec-

molecule with and MBNL1 protein is —541.58 a.u. The molecular trophilic center. Full Fitness score corresponding to docking of the docking of title molecule suggests that it can bind and inhibit the

receptor enzymes.

Appendix A. Supplementary data

Supplementary data to this article can be found online at <https://doi.org/10.1016/j.ejbas.2018.10.001>.

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