

# Atomic Physics

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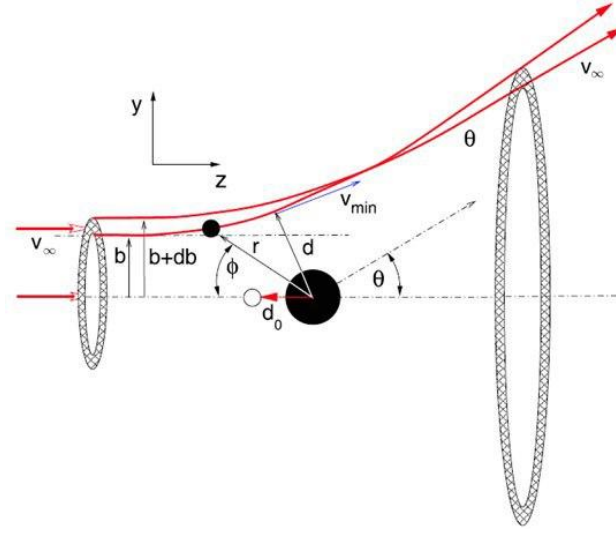
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## Chapter 1

# Rutherford's Alpha Particle Scattering Experiment



According to Coulomb's Law:

$$F = \frac{1}{4\pi\epsilon_0} \cdot \frac{Z_1 Z_2 e^2}{r^2} = \frac{C}{r^2}$$

$$C = \frac{Z_1 Z_2 e^2}{4\pi\epsilon_0}$$

$$F_y = F \sin \phi = \frac{C}{r^2} \sin \phi$$

Law of momentum and angular momentum

$$mv_y = \int F_y dt$$

$$mr^2 \dot{\phi} = mv_\infty b$$

Then we integrate

$$\begin{aligned} v_y &= \frac{1}{m} \int \frac{C}{r^2} \sin \phi dt = \frac{1}{m} \int \frac{C}{r^2} \sin \phi \frac{dt}{d\phi} d\phi = \frac{1}{m} \int \frac{C}{r^2} \sin \phi \frac{r^2}{v_\infty b} d\phi = \frac{C}{mv_\infty b} \int_0^{\pi-\theta} \sin \phi d\phi \\ &= \frac{C}{mv_\infty b} (1 + \cos \theta) \end{aligned}$$

Now we need to relate  $\theta$  with  $b$ , Since

$$v_y = v_\infty \sin \theta$$

We have

$$\frac{C}{mv_\infty b} (1 + \cos \theta) = v_\infty \sin \theta$$

So that

$$\cot \frac{\theta}{2} = \frac{1 + \cos \theta}{\sin \theta} = \frac{mv_\infty^2 b}{C} = \frac{2E_0 b}{C}$$

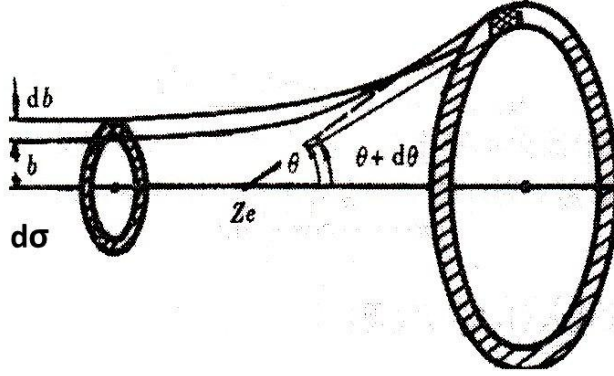
Note that this trigonometry transform is used

$$\frac{1 + \cos \theta}{\sin \theta} = \frac{2 \cos^2 \frac{\theta}{2}}{2 \sin \frac{\theta}{2} \cos \frac{\theta}{2}} = \cot \frac{\theta}{2}$$

Finally

$$b = \frac{C}{2E_0} \cdot \cot \frac{\theta}{2} = \frac{1}{4\pi\epsilon_0} \frac{Z_1 Z_2 e^2}{2E_0} = \frac{1}{4\pi\epsilon_0} \cdot \frac{Z_1 Z_2 e^2}{mv_\infty^2} \cdot \cot \frac{\theta}{2}$$

Now we begin to find the relation between  $db$  and  $d\Omega$



$$d\sigma = 2\pi b db = \pi \left( \frac{1}{4\pi\epsilon_0} \right)^2 \left( \frac{Z_1 Z_2 e^2}{mv_\infty^2} \right)^2 \frac{\cos \frac{\theta}{2}}{\sin^3 \frac{\theta}{2}} d\theta$$

$$\Omega = 2\pi (1 - \cos \theta)$$

$$d\Omega = 2\pi \sin \theta d\theta = 4\pi \sin \frac{\theta}{2} \cos \frac{\theta}{2} d\theta$$

Then we found  $\frac{d\sigma}{d\Omega}$  is only related with  $\theta$

$$\frac{d\sigma}{d\Omega} = \left( \frac{1}{4\pi\epsilon_0} \right)^2 \left( \frac{Z_1 Z_2 e^2}{2mv_\infty^2} \right)^2 \sin^{-4} \frac{\theta}{2}$$

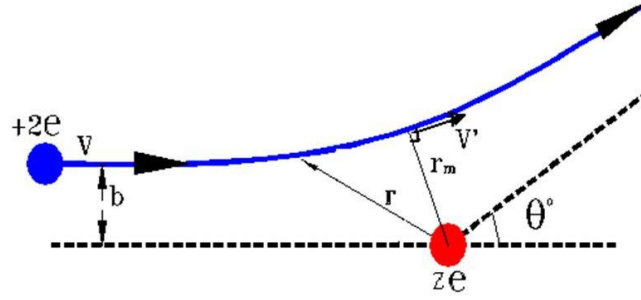
As for a thin gold leaf, we assume there's only one layer of atoms, the density of atoms is  $N$ , the area and thickness of gold leaf are  $A$  and  $t$

When  $n$  particles passed through the gold leaf,  $dn$  of them ended up in  $d\Omega$

$$\frac{dn}{n} = \frac{NAt d\sigma}{A} = Nt d\sigma = Nt \left( \frac{1}{4\pi\epsilon_0} \right)^2 \left( \frac{Z_1 Z_2 e^2}{2mv_\infty^2} \right)^2 \sin^{-4} \frac{\theta}{2} d\Omega$$

$$\frac{dn}{d\Omega} \sin^4 \frac{\theta}{2} = nNt \left( \frac{1}{4\pi\epsilon_0} \right)^2 \left( \frac{Z_1 Z_2 e^2}{2mv_\infty^2} \right)^2 = \text{const}$$

For alpha particles,  $Z_1 = 2$ . We can also take the closest distance between the 2 particles as the radius of a particle.



What we have already known are:

$$\frac{1}{2}MV^2 = \frac{1}{2}MV'^2 + \frac{2Ze^2}{4\pi\epsilon_0 r_m}$$

$$MVb = MV'r_m$$

$$b = \frac{1}{4\pi\epsilon_0} \cdot \frac{2Ze^2}{MV} \cdot \cot \frac{\theta}{2}$$

We solve  $r_m$  with the equation we known.

$$\frac{1}{4\pi\epsilon_0} \cdot \frac{2Ze^2}{M} \cdot \cot \frac{\theta}{2} = V'r_m$$

$$\frac{1}{2}MV^2 = \frac{1}{2}M \left( \frac{1}{4\pi\epsilon_0} \cdot \frac{2Ze^2}{Mr_m} \cdot \cot \frac{\theta}{2} \right)^2 + \frac{2Ze^2}{4\pi\epsilon_0 r_m}$$

$$\frac{1}{2}MV^2 r_m^2 - \frac{2Ze^2}{4\pi\epsilon_0} r_m - \frac{1}{2}M \left( \frac{1}{4\pi\epsilon_0} \cdot \frac{2Ze^2}{M} \cdot \cot \frac{\theta}{2} \right)^2 = 0$$

Finally

$$r_m = \frac{1}{4\pi\epsilon_0} \frac{2ze^2}{MV^2} \left( 1 + \frac{1}{\sin \frac{\theta}{2}} \right)$$



## Chapter 2

# The Energy and Radiation of Atoms

### 2.1 Rydberg Constant and Wavelength of Radiation

Balmer Formula for Hydrogen Atoms:

$$\frac{1}{\lambda} = R_H \left[ \frac{1}{m^2} - \frac{1}{n^2} \right]$$

Let

$$T(m) = \frac{R_H}{m^2}$$
$$T(n) = \frac{R_H}{n^2}$$

Then

$$\frac{hc}{\lambda} = hc [T(m) - T(n)] = \frac{R_H hc}{m^2} - \frac{R_H hc}{n^2}$$

### 2.2 Bohr's Theory of Hydrogen Atoms

Energy of steady states:

$$E_n = -\frac{1}{2} \frac{Ze^2}{4\pi\epsilon_0 r_n}$$

Transition

$$h\nu = E_n - E_m$$

Angular Momentum

$$L = n \cdot \frac{h}{2\pi} = n\hbar = m_e v_n r_n$$
$$m_e \frac{v_n^2}{r_n} = \frac{Ze^2}{4\pi\epsilon_0 r_n^2}$$

From which we can indicate that the radius is qumntumized

$$\frac{n^2 \hbar^2}{m_e r_n^3} = \frac{Ze^2}{4\pi\epsilon_0 r_n^2}$$
$$r_n = \frac{4\pi\epsilon_0 \hbar^2}{m_e e^2} \cdot \frac{n^2}{Z} = a_0 \cdot \frac{n^2}{Z}$$

For hydrogen atoms,  $Z = 1$ . The energy is also qumntumized.

$$E_n = -\frac{1}{2} \frac{Ze^2}{4\pi\epsilon_0} \frac{m_e e^2}{4\pi\epsilon_0 \hbar^2} \cdot \frac{Z}{n^2} = -\frac{m_e e^4}{2(4\pi\epsilon_0 \hbar)^2} \cdot \frac{Z^2}{n^2}$$

The velocity is also quantized

$$v_n = \frac{n\hbar}{m_e r_n} = \frac{n\hbar}{m_e} \frac{m_e e^2}{4\pi\epsilon_0 \hbar^2} \cdot \frac{Z}{n^2} = \frac{e^2}{4\pi\epsilon_0 \hbar} \cdot \frac{Z}{n} = \frac{Z\alpha c}{n}$$

$$\alpha = \frac{e^2}{4\pi\epsilon_0 \hbar c} = \frac{1}{137}$$

Calculate the value of Rydberg Constant

$$E = -\frac{m_e e^4}{2(4\pi\epsilon_0 \hbar)^2} \cdot \frac{Z^2}{n^2} = -\frac{2\pi^2 m_e e^4}{(4\pi\epsilon_0 \hbar)^2} \cdot \frac{Z^2}{n^2}$$

$$\frac{hc}{\lambda} = E_2 - E_1 = \frac{2\pi^2 m_e e^4 Z^2}{(4\pi\epsilon_0 \hbar)^2} \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$

$$\frac{1}{\lambda} = \frac{2\pi^2 m_e e^4 Z^2}{(4\pi\epsilon_0)^2 \hbar^3 c} \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$

$$R = \frac{2\pi^2 m_e e^4 Z^2}{(4\pi\epsilon_0)^2 \hbar^3 c}$$

For hydrogen atoms

$$R_H = \frac{2\pi^2 m_e e^4}{(4\pi\epsilon_0)^2 \hbar^3 c}$$

## 2.3 Rydberg Constant of Different Atoms

In previous sections, we assumed that the nucleus is fixed at a point, with the electron surrounded. But in fact, the nucleus's mass is not infinity, and it moves as well. So that for different atoms, the Rydberg Constants varies. In a two-body system, we define the Reduced Mass of a system as

$$\mu = \frac{Mm}{M+m}$$

And we replace the Reduced Mass with  $m_e$

$$R_\infty = \frac{2\pi^2 m_e e^4}{(4\pi\epsilon_0)^2 \hbar^3 c}$$

$$R_A = R_\infty \cdot \frac{1}{1 + \frac{m}{M}}$$