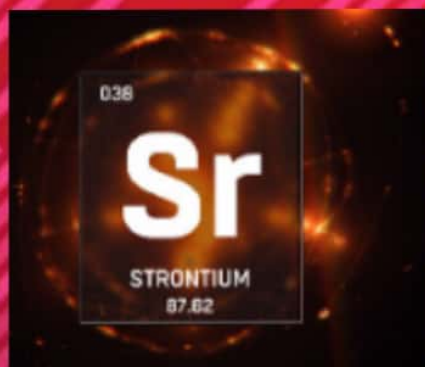


DAILY LIFE USAGE :-

1. Used to make glass for color tv
2. It is incorporated in the bones.
3. $^{87}\text{Sr}/^{86}\text{Sr}$ used to determine provenance areas of sediment.
4. Strontium aluminate is used in "glow in dark toys".
5. Strontium carbonate is used in fireworks to give deep red colour.
6. Strontium carbonate is used in manufacture of ferrite magnets.
7. Strontium chloride is used in production of toothpaste for sensitive teeth.
8. Radioactive strontium is used as an active ingredient in "Metastron" which is used as radiopharmaceutical for bone pain.
9. ^{90}Sr is used as power source for "radioisotope thermoelectric generators".



Birth/Discovery:-
Strontianite after
Stronian a town
in Scotland.
Discovered by
William Cruickshank.



POSITION IN PERIODIC TABLE:-
PERIOD 5
AND GROUP 2,
ALKALINE EARTH METAL.

*Isotopes: the stable
isotopes of Strontium
found in nature are*

^{84}Sr
 ^{86}Sr
 ^{87}Sr
 ^{88}Sr
 $^{90}\text{Sr}^*$

**Chemical and physical
properties:-**

1. Phase at STP:- Solid
2. Melting point:- 777°C
3. Boiling Point:- 1377°C
4. Density:- 2.64 g/cm^3
5. Oxidation state:- +1, +2
6. Atomic radius:- 215 pm
7. Paramagnetic nature
8. Weight:- 87.62 g
9. Configuration:- $[\text{Kr}]5s^2$



SPECIAL COMPOUNDS

STRONTIUM ALUMINATE (
 SrAl_2O_4), STRONTIUM
CARBONATE (SrCO_3), STRONTIUM
CHROMATE (SrCrO_4), STRONTIUM
SULFATE (SrSO_4).

Uniqueness quotient:-
When Sr is mixed with
paint, it causes paint
to glow in dark.