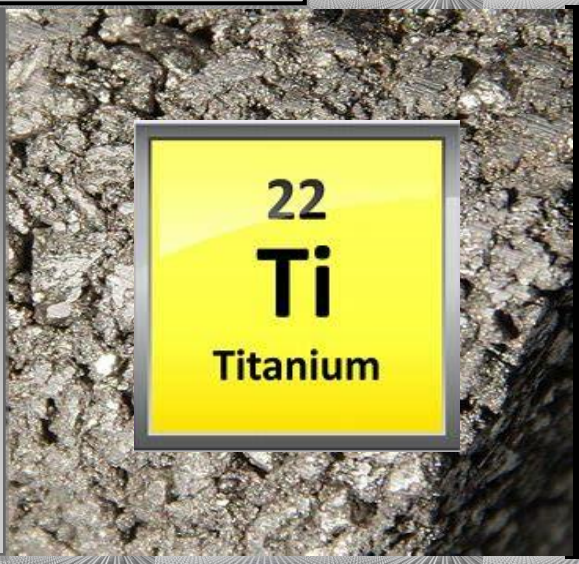
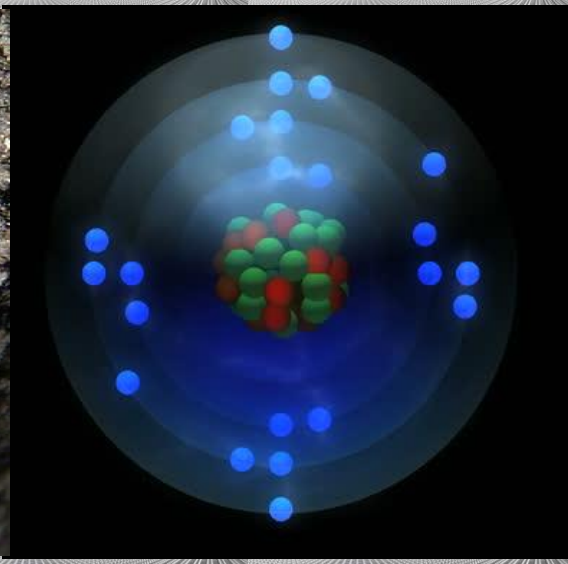
Facts

- ⇒ **Symbol: Ti**
- ⇒ **Atomic Number: 22**
- ⇒ **Classification: Transitional Metal**
- ⇒ **Atomic Mass : 47.867**
- ⇒ **Phase at room temperature: Solid**
- ⇒ **Density: 4.506/cm^3**
- ⇒ **Melting Point: 1668°C, 3034°F**
- ⇒ **Boiling Point: 3287°C, 5949°F**
- ⇒ **Strong and light ,relatively low electrical and thermal conductivity**
- ⇒ **Inactive and is very resistant to corrosion from other elements and acids and O_2.**

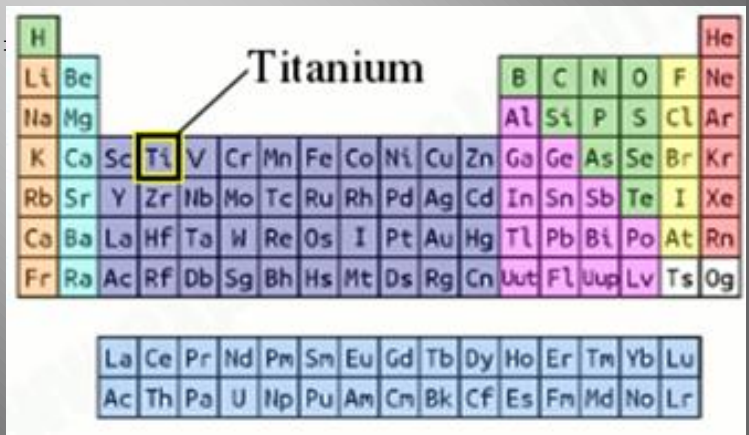
Origin And History

⇒ **Titanium was first recognized as a new element by Reverend William Gregor in 1791.He named the element enachanite. The name was later changed to titanium by German chemist M.H. Kalproth. The first pure titanium was produced by American chemist M. A. Hunter in 1910.**





Position in Periodic Table



Titanium is the positioned in group 4, 4th period and d block of the periodic table.

Isotopes

⇒ **Titanium has five stable isotopes including titanium-46, 47, 48, 49, and 50. The majority of titanium found in nature is in the form of the isotope titanium-48.**

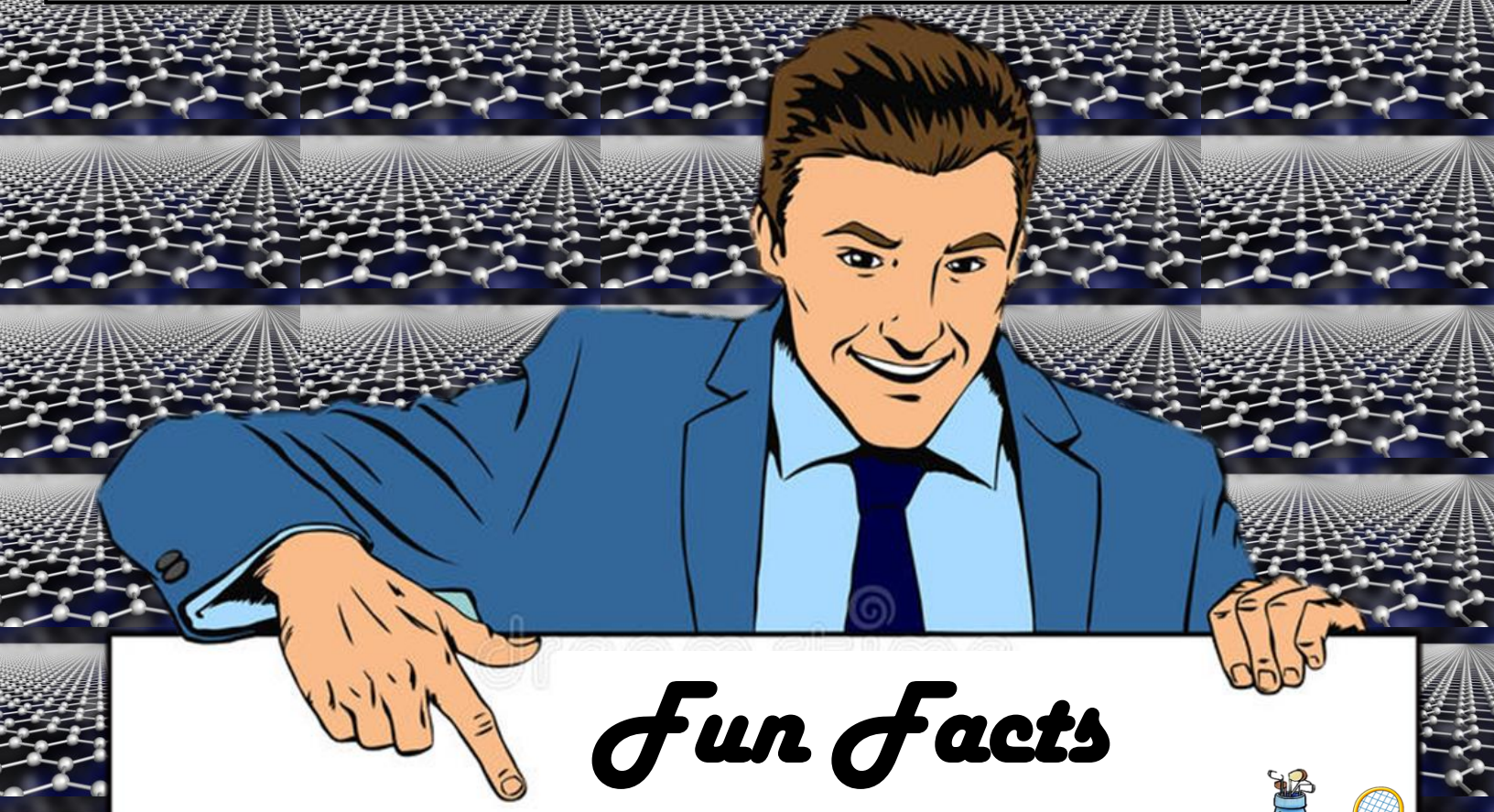
Isotope	Abundance	Mass (amu)
^{46}Ti	71.500%	45.95263
^{48}Ti	17.500%	47.94795
^{50}Ti	11.000%	49.94479

Uses

⇒ Used to form titanium dioxide (TiO_2) which number of industrial uses including white paint, paper, plastics, and cements.

⇒ Used to alloy with different metals such as iron, aluminum, and manganese where it helps to produce strong and lightweight alloys for use in spacecraft, naval ships, missiles, and as armor plating.

⇒ used in various medical applications such as hip replacements and dental implants and also used in jewelry to make rings and watches.



Fun Facts

⇒ Titanium oxide is often used with graphite to make high-end golf clubs and tennis rackets.

⇒ It is the only element that will burn in pure nitrogen gas

⇒ It is found in meteorites, on the Moon, and in some types of stars.

⇒ The Guggenheim Museum in Bilbao, Spain is covered with titanium plated tiles.