Zinc (Zn)

Atomic number: 30

Atomic mass: 65.38

Description: The name is derived from the German, ‘zinc’, which may in turn be derived from the Persian word ‘sing’, meaning stone.

Discovered by: Andreas Marggraf in 1746 but known to the Greeks and Romans before 20BC.

Properties:

Melting point: 692.677K Boiling Point: 1180K

Density: 7.134 g/cm^3 State: Solid

Group: 12 Period: 4 Block: d

Electronic Configuration: [Ar] $3d^{10}4s^2$

Isotopes: ^{64}Zn

Appearance: A silvery-white metal with a blue tinge. It tarnishes in air.

Uses:

1) Most zinc is used to galvanize other metals, such as iron, to prevent rusting. Galvanized steel is used in car bodies, street lamp post, safety barriers and suspension bridges.

2)Large quantities of zinc are used to produce die-casting, which are important in the automobile, electrical and hardware industries. Zinc is also used in alloys such as brass, nickel silver and aluminum solder.

3) Zinc oxide is widely use in the manufacture of very many products such as paints, rubber, cosmetics, pharmaceuticals, plastics, inks, soaps, batteries, textile and other electrical equipment.

4) Zinc sulphide is also used in making luminous paints, florescent lights and x-rays screens.

Source: www.rsc.org