Molecular polarity

Friday, 16 February 2024 10:51 am

Non - metal + nonmetal - simple covalent

Cyclohexane is non-polar

Electronegativity (idea from the simple covalent)

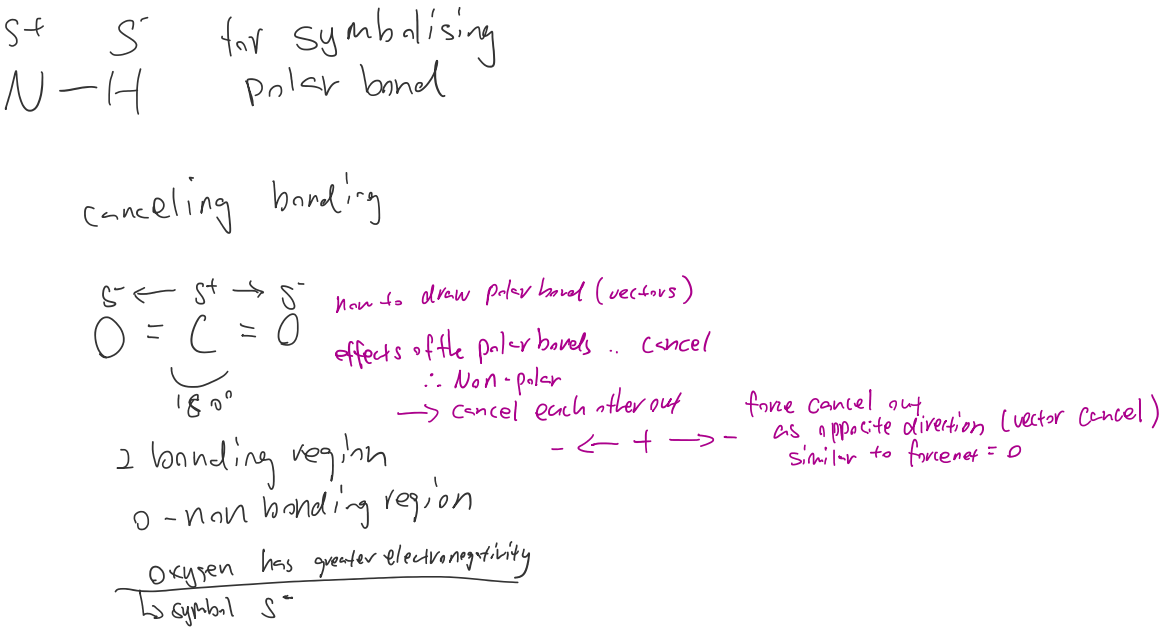
- Attraction for a pair of electron

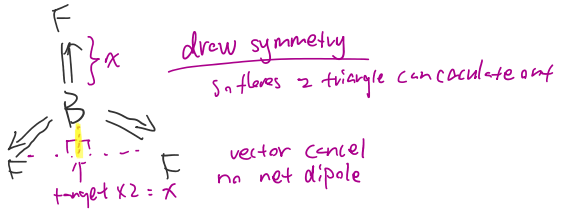
If the bond is non - polar then the molecule is non polar, there is exception
If the molecule have polar bond, molecule could be polar as long as they are not cancel out

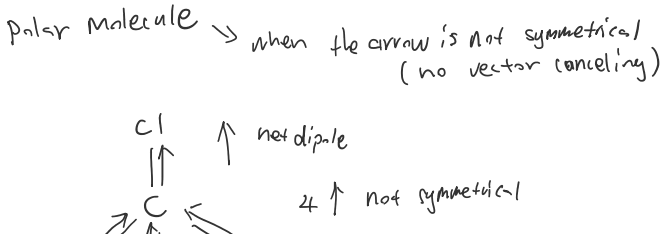
Polar bond cancelling polar bond in molecule

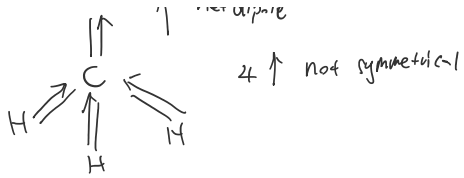
Decide whether the molecule are polar - see whether there is vector canceling (polar cancelling) (or there is symmetry) (sometime the molecule can be polar with non polar end which the overall is polar)

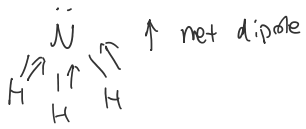
- An asymmetrical molecule with polar bonds will have an overall net dipole movement (eg: positive end and negative end) (vector chanceless)








4 ↑ not symmetrical


↑ net dipole