

Ion and preview

Sunday, 28 January 2024 1:15 pm

Cation +

Anion -

Type of structure

- Ionic
 - o Metal + non metal
- Metal
 - o Metal
- Covalent(Giant covalent & Simple covalent)
 - o Nonmetal bonding

Ionic bond - electrostatic attraction between two opposite charge ion

Lattice

Lattice structure - regular repeating pattern

(metal ion arranged in regular repeating pattern)

Lattice enthalpy - the energy require to break (1mole) of an ionic lattice into its individual constituent ions.

Solubility:

When ionic compound dissolve in water, the polar water molecule surround and separate the individual ion due to their attraction to the opposite charge.

Transition metal - set of metallic element occupying the central block in periodic table

Polyatomic ion - ion composed of more than one atom.

Lattice structure - regular repeating pattern

Reglar repeating pattern of cation and anion

Regular repeating pattern of cation and electron

Lattice enthalpy - the energy require to break 1 mole of ionic lattice into its individual constituent ions

The energy require to break 1 mole of ionic lattice into individual constituent ion

Regular repeating pattern of ions