* + Unit 1:
    1. One liter is equal to:
       1. 1,000,000 picoliters
       2. 1,000,000,000 picoliters
       3. **1,000,000,000,000 picoliters**
       4. 1,000,000,000,000,000 picoliters
    2. The unit symbol for micrometers is:
       1. mm
       2. **µm**
       3. mmm
       4. ŋm
    3. One gigameter is equal to:
       1. 1,000,000 meters
       2. **1,000,000,000 meters**
       3. 1,000,000,000,000 meters
       4. 1,000,000,000,000,000 meters
    4. One centigram is equal to:
       1. 0.1 milligrams
       2. 1 milligram
       3. **10 milligrams**
       4. 100 milligrams
    5. An ionic bond is made of:
       1. **A metal and a nonmetal**
       2. Two or more metals
       3. A metal and at least two nonmetals
       4. Two or more nonmetals
    6. The Tyndall effect is useful for detecting:
       1. Mixtures
       2. Suspensions
       3. Solutions
       4. **Colloids**
    7. A teaspoon of salt is dissolved into a cup of water. Which is the solute, and which is the solvent?
       1. The salt is the solvent and the water is the solute
       2. **The salt is the solute and the water is the solvent**
       3. Both are solutes
       4. Both are solvents
    8. The chemical formula for dichromate is:
       1. **Cr2O7--**
       2. Cr2O8--
       3. Cr2O7-
       4. Cr2O8-
  + Unit 2:
    1. \_\_\_\_\_\_\_\_\_ is an extensive property.
       1. The copper in the test tube reacts with oxygen
       2. The copper in the test tube is blue
       3. **There are 10 grams of copper in the test tube**
       4. The copper in the test tube is heavy
    2. \_\_\_\_\_\_\_\_\_ is a chemical property.
       1. Water is clear
       2. Water freezes at 0oC
       3. Water is wet
       4. **Water and cesium explode when exposed to each other**
    3. \_\_\_\_\_\_\_\_\_ are fluids.
       1. Solids
       2. Liquids
       3. Gasses
       4. **Both B & C**
    4. \_\_\_\_\_\_\_\_\_ do not have definite volumes.
       1. Solids
       2. Liquids
       3. **Gasses**
       4. Both B & C
    5. \_\_\_\_\_\_\_\_\_ have definite shapes.
       1. **Solids**
       2. Liquids
       3. Gasses
       4. All of the above
    6. The chemical formula for Lead (IV) Fluoride is:
       1. PbF
       2. **PbF4**
       3. PbF2
       4. Pb2F2
    7. The proper name for N2O4 is:
       1. Nitrogen Quadroxide
       2. Nitrogen Tetroxide
       3. Dinitrogen Quadroxide
       4. **Dinitrogen Tetroxide**
    8. The prefix "Hepta" means:
       1. 5
       2. 6
       3. **7**
       4. 8
  + Unit 3:
    1. The significant figures in 3000 are:
       1. **3000**
       2. 3000
       3. 3000
       4. 3000
    2. The significant figures in 0300.0 are:
       1. 0330.0
       2. 0300.0
       3. **0300.0**
       4. 0300.0
    3. 4.2 + 3.57 =
       1. 7.77
       2. 8.0
       3. **7.8**
       4. 8
    4. 7.6 x 18.87 =
       1. 143.412
       2. 143.4
       3. 143
       4. **140**
    5. The formula for density is:
       1. **Mass/Volume**
       2. Volume/Mass
       3. Volume x 2/Mass
       4. None of the above
    6. Which is the proper format for dimensional analysis? (2 liters is how may milliliters)
       1. **?mL = 2L**
       2. 2L = ?mL
       3. mL → 2L
       4. 2L → mL
    7. How many grams are in 54439.943 milligrams?
       1. 54
       2. 54.44
       3. 54.440
       4. **54.439943**
    8. How many milliliters are in 4.5 cubic centimeters?
       1. 0.45
       2. **4.5**
       3. 0.5
       4. 5.0
  + Unit 4:
    1. \_\_\_\_\_\_\_\_ cannot exist as a single atom.
       1. Oxygen
       2. Iodine
       3. Bromine
       4. **All of the above**
    2. The formula for nitrogen gas is:
       1. N
       2. **N2**
       3. N3
       4. N4
    3. Balance this equation: Sodium plus Chlorine yield Sodium Chloride.
       1. Na + Cl → NaCl
       2. Na2 + Cl2 → Na2Cl2
       3. Na2 + 2Cl → 2NaCl
       4. **None of the above**
    4. Balance this equation: Lead plus Chlorine yield Lead (IV) Chloride.
       1. Pb + Cl2 → PbCl2
       2. Pb + Cl4 → PbCl4
       3. **Pb + 2Cl2 → PbCl4**
       4. Pb + 4Cl → PbCl4
    5. Balance this equation: Ozone yields Oxygen.
       1. O3 → O2
       2. 2O → O2
       3. O4 → 2O2
       4. **2O3 → 3O2**
    6. What type of chemical reaction is #4 (Lead plus Chlorine yield Lead (IV) Chloride)?
       1. **Synthesis/Composition**
       2. Decomposition/Combustion
       3. Single Replacement
       4. Double Replacement
    7. A substance with a pH of 10 is:
       1. Acidic
       2. **Basic**
       3. Neutral
       4. You can't tell
    8. How do acids interact with Hydrogen? (H+)
       1. **They donate it**
       2. They receive it
       3. They destroy it
       4. We don't know
  + Unit 5:
    1. A mole is equal to how much of a substance?
       1. **6.02 x 1023**
       2. 6.02 x 1024
       3. 6.02 x 1025
       4. 6.02
    2. The Kinetic Molecular Theory perfectly describes:
       1. All gasses
       2. Gasses at STP
       3. Gasses at absolute zero
       4. **Ideal gasses**
    3. Gasses are most ideal when:
       1. Pressure and temperature are both low
       2. Pressure and temperature are both high
       3. **Pressure is low and temperature is high**
       4. Pressure is high and temperature is low
    4. \_\_\_\_\_\_\_\_\_\_ is the most useful unit for making calculations with gasses.
       1. Celsius
       2. **Kelvin**
       3. Fahrenheit
       4. All are equally useful
    5. Pressure can be best described as:
       1. Force
       2. Force per volume
       3. **Force per area**
       4. Force per mass
    6. If the temperature of a gas increases and the pressure stays the same, the volume of the gas:
       1. **Increases**
       2. Decreases
       3. Stays the same
       4. It's random
    7. Question #6 describes what law? (Temperature's relation to Pressure)
       1. **Charles's Law**
       2. Boyle's Law
       3. Bernoulli's Principle
       4. Torricelli's Law
    8. If the temperature of a gas increases and the volume stays the same, the pressure:
       1. **Increases**
       2. Decreases
       3. Stays the same
       4. We don’t know