Lab 8

	Trial 1	Trial 2
KHP (g)	.063	.06
Initial Volume (mL)	.2	.13
Final Volume (mL)	33.8	31.4

mm maleic acid = 116.1 g

$.006 \cdot \dfrac{50}{1000} \cdot 116.1 = .013483$ g ≈ .035 g exact

$pH = pKa + \log\left(\dfrac{C_4H_3O_4^-}{C_4H_4O_4}\right)$ $pH = pKa + \log\left(\dfrac{C_4H_2O_4^{-2}}{C_4H_3O_4^{-1}}\right)$

0   0