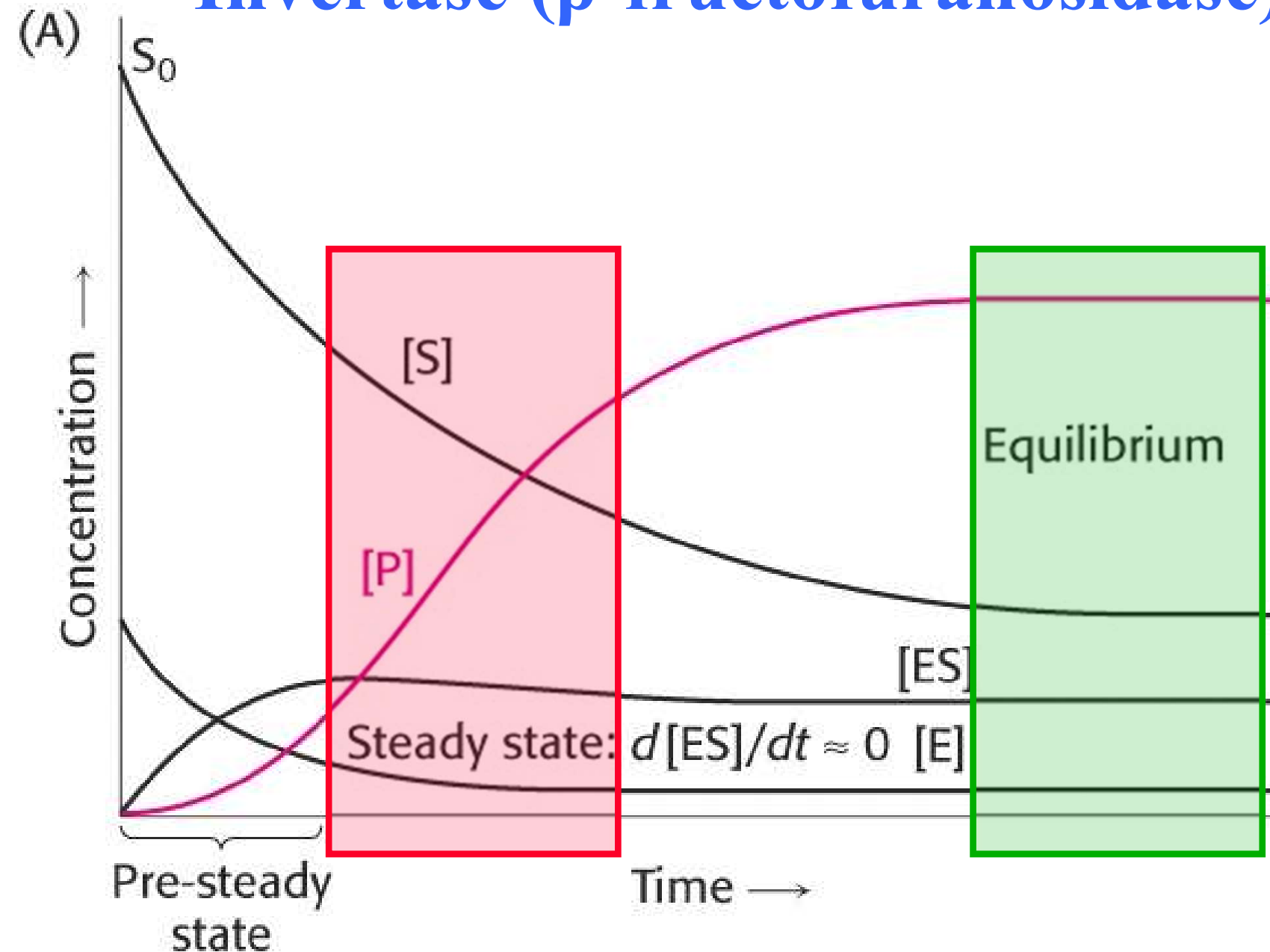
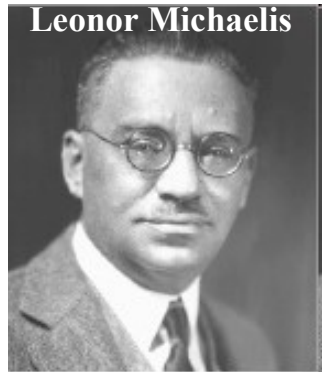


Applying our knowledge to enzymes: Invertase (β -fructofuranosidase)



What about (common) reversible chemistry?



$$\Rightarrow v = \frac{\frac{V_{\max}^f [\text{S}]}{K_M^S} - \frac{V_{\max}^r [\text{P}]}{K_M^P}}{1 + \frac{[\text{S}]}{K_M^S} + \frac{[\text{P}]}{K_M^P}}$$

And for initial $v_0 \Rightarrow [\text{P}] = 0$:
Michaelis Menten!

$$V_{\max}^f = k_2 [\text{E}]_{\text{T}} \quad V_{\max}^r = k_{-1} [\text{E}]_{\text{T}}$$

$$K_M^S = \frac{k_{-1} + k_2}{k_1} \quad K_M^P = \frac{k_{-1} + k_2}{k_{-2}}$$

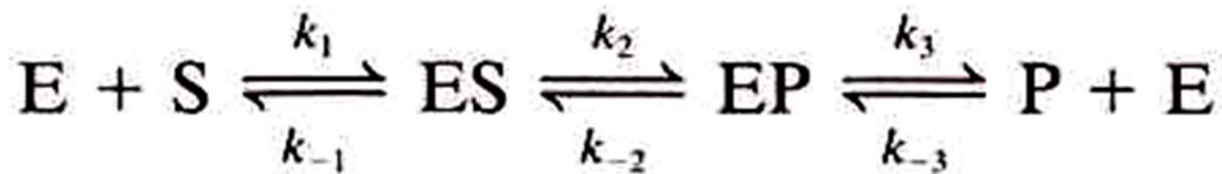
$$[\text{E}]_{\text{T}} = [\text{E}] + [\text{ES}]$$

@ equilibrium: $v = 0$

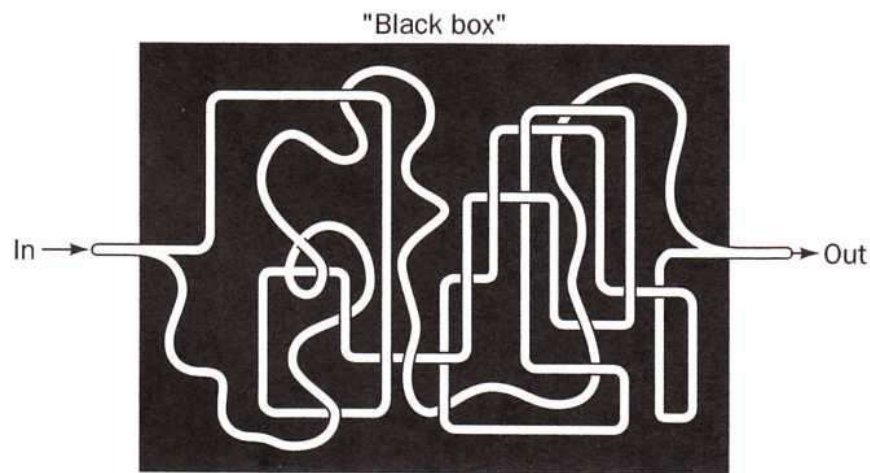
$$K_{\text{eq}} = \frac{[\text{P}]_{\text{eq}}}{[\text{S}]_{\text{eq}}} = \frac{V_{\max}^f K_M^P}{V_{\max}^r K_M^S}$$

Haldane relationship

What if there are still more reversible steps?



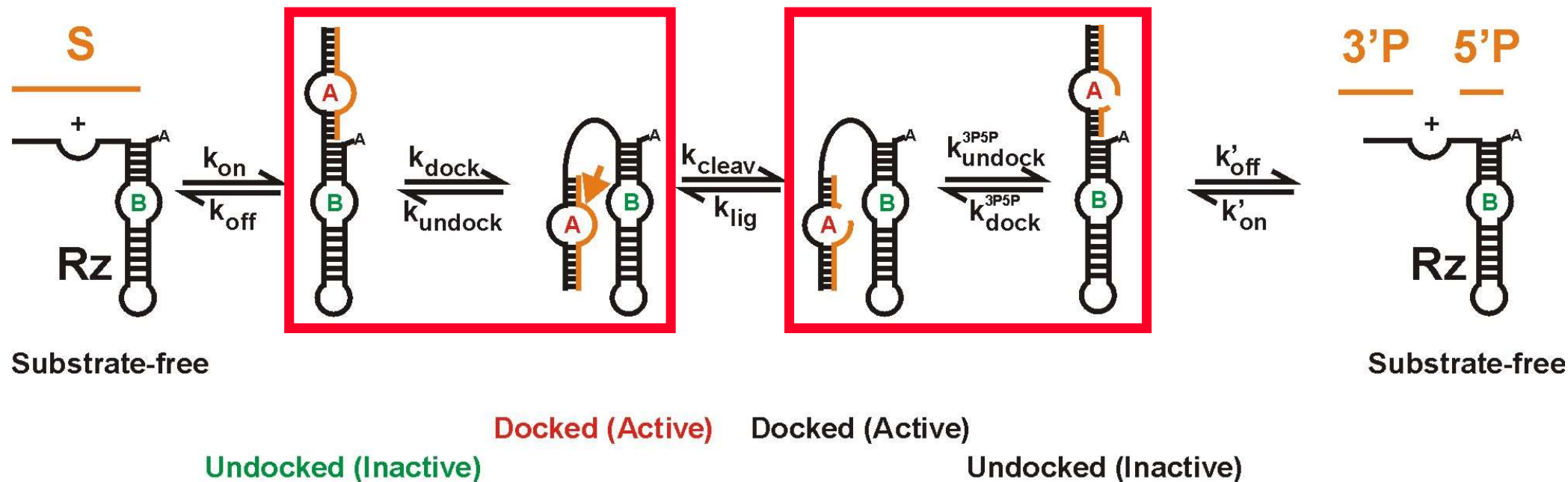
While $[\text{S}]_0$ and $[\text{P}]_0$ can be manipulated under steady-state conditions, the resulting equations do not allow to solve for all six (or even more) unknown constants
 \Rightarrow we have to look into the black box!



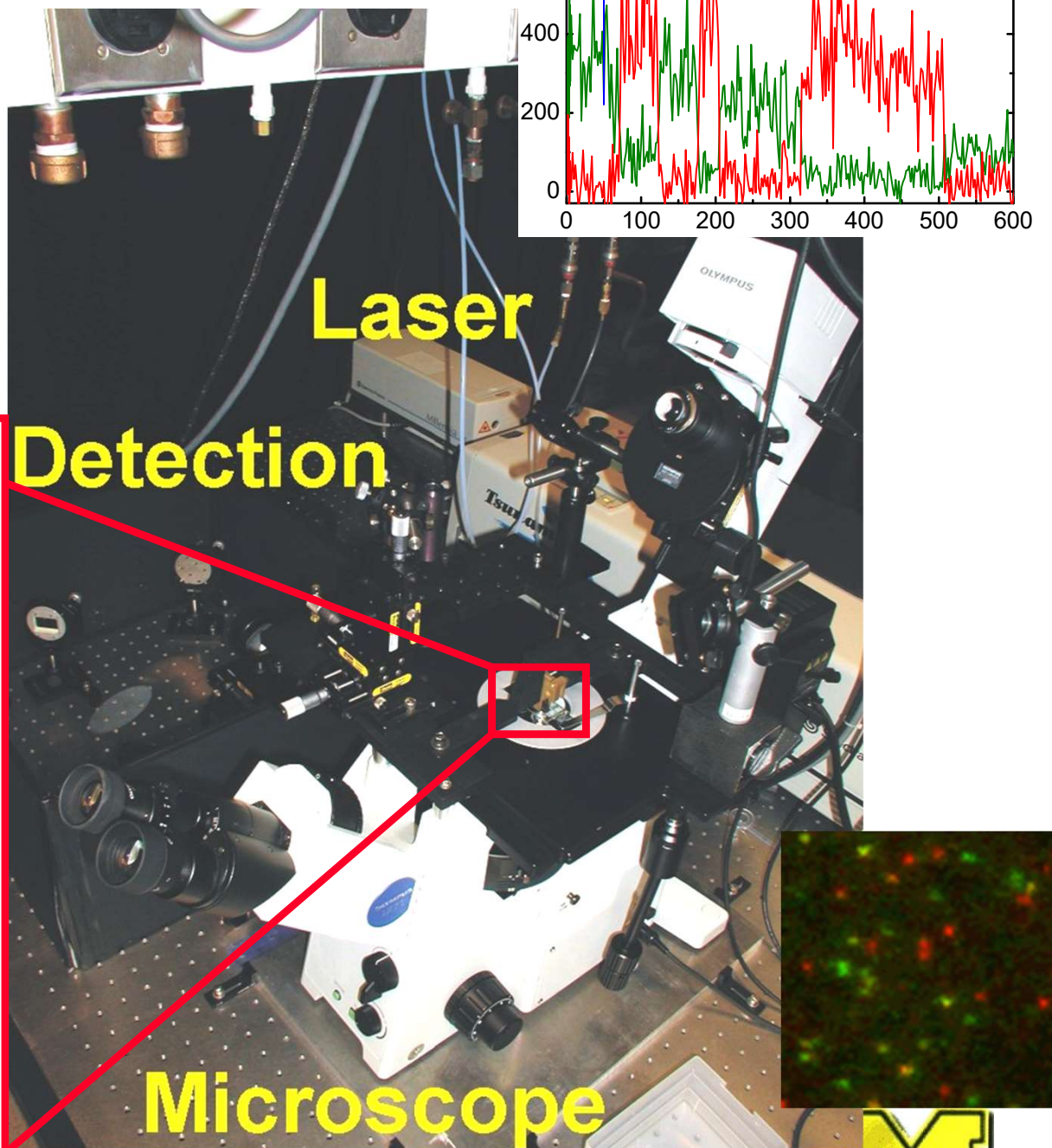
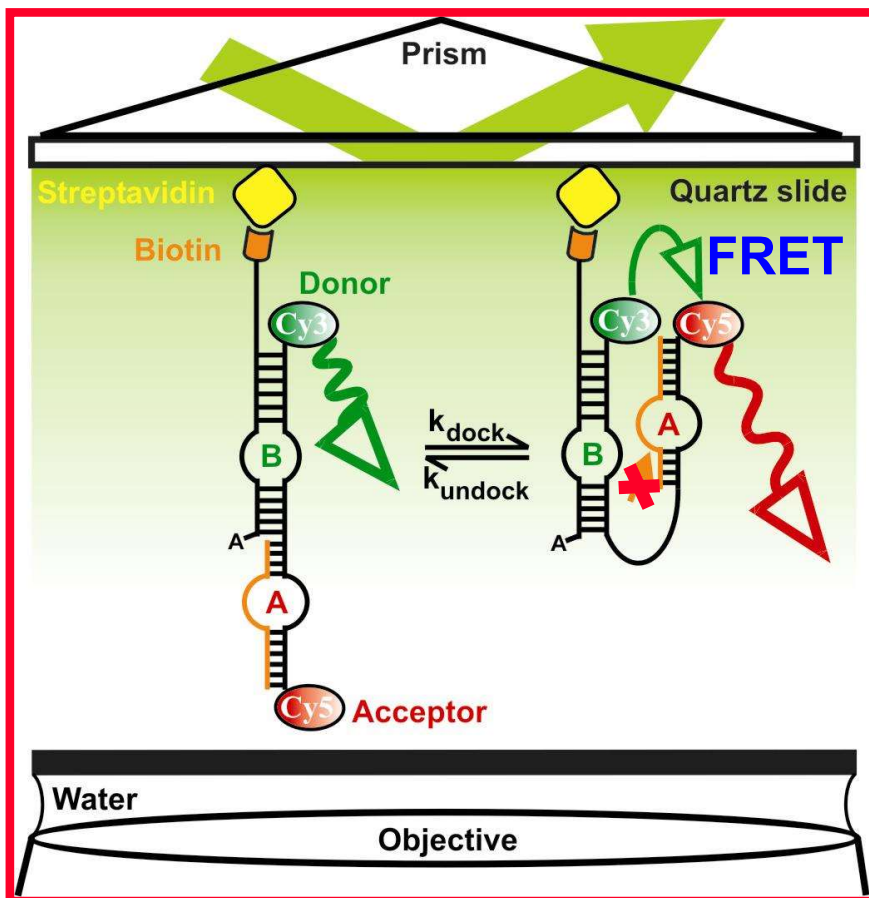
In addition, steady-state kinetic measurements:

- only yield phenomenology, not mechanism
- cannot unambiguously establish a mechanism
- can only rule out some alternative mechanisms

Case in point -- The hairpin ribozyme: Reversible domain docking is a crucial step on the reaction pathway



We measured docking rate constants by single molecule FRET



Nils Walter: Chem 451

Zhuang, X., Kim, H., Pereira, M.J.B., Babcock, H.P., Walter, N.G. and Chu, S. *Science* 296 (2002) 1473-1476



Kinetic modeling then allowed us to extract the unobservable reversible chemistry rate constants

