

Every underlined topic is a leading node and the “+” labelled topics are the sub-nodes in a “consists of” relationship to their respective leading node. The “-” labelled sub points are the sub nodes of their respective “+” sub nodes. Remove duplicate nodes. Each node is to come only once

Introduction to Organic Chemistry

- + Description: Study of carbon-containing compounds and their properties

Nomenclature

- + IUPAC Rules

- Rules for naming organic compounds
- 3 rules listed (e.g., names of complex substituents are enclosed in parentheses)

- + Functional Groups

- Definition: Atoms or groups of atoms that determine the chemical properties of an organic compound
- Examples: -OH (Hydroxyl), -COOH (Carboxyl), -CHO (Aldehyde), -CO- (Ketone)

Isomerism

- + Structural Isomerism

- Definition: Compounds with the same molecular formula but different structures
- Types: Chain Isomerism, Position Isomerism, Functional Isomerism, Metamerism

- + Stereoisomerism

- Definition: Compounds with the same molecular formula and bond order, but differing in 3D arrangement of atoms in space
- Types: Geometric Isomerism, Optical Isomerism

Chemical Bonding

- + Sigma (σ) Bond

- Definition: A covalent bond formed by end-to-end overlap of atomic orbitals

- + Pi (π) Bond

- Definition: A covalent bond formed by lateral overlap of atomic orbitals

- + Hybridization

- Definition: The process of mixing atomic orbitals to form new hybrid orbitals suitable for bonding

- Types: sp^3 , sp^2 , sp

Inductive and Electromeric Effects

+ Inductive Effect (I effect)

- Definition: A permanent polarity effect transmission through a σ -bond
- Types: +I effect, -I effect

+ Electromeric Effect (E effect)

- Definition: A temporary polarity effect transmission through a π -bond
- Types: +E effect, -E effect