

## *Answer Key Sweet 16 Chemistry Compound Tournament*

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Sweet 16 Chemistry Ion Tournament Publication No. 10510 With spring just around the corner, your students' thoughts will soon be turning to sunshine, prom, and the NCAA basketball tournament. This clever activity combines the ever-popular basketball office pool with a review of the formulas of ionic compounds and their solubility.

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### **Sweet 16 Chemistry Compound Tournament Answer Key**

Sweet 16 Chemistry Ion Tournament First Round Second Round Semifinals Finals Winner  $\text{NiCl}_2$   $\text{Ni(II)}_2 + \text{Fe(II)}$   $\text{S}$   $\text{Zn}$   $\text{CO}_3$   $\text{Na}$   $\text{CrO}_4$   $\text{Mg}$   $\text{PO}_4$   $\text{Ag(I)}$   $\text{SO}_4$   $\text{Ba}$   $\text{NO}_3$   $\text{Al}$   $\text{OH}$  Rules 1. First Round →Add Charges 2. Second Round →Write Formula 3. Semis →Soluble Compound (compound that will dissolve in water) 4. Finals →Soluble Product of Chemical Reaction 5.

### **Name Class/Period Sweet 16 Chemistry Ion Tournament**

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the structure and charges of ions, and molar mass calculations. The result is the Sweet 16 Chemistry Compound Tournament. Hopefully, your students will enjoy playing the "Tournament" while reviewing these important topics. The rules for filling out the Tournament brackets are simple: For the first round, the student determines the chemical ...

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First round: Predict the Gaseous product. Second round: Predict the one that dissolves in water to

give an acidic solution. Semis: The gas that is more dense. Finals: The winner is a greenhouse gas produced in volcanic emissions. First round: Sodium Carbonate vs. Hydrochloric acid Zinc vs. Hydrochloric acid Hydrogen peroxide vs. Yeast (catalyst) Sodium Sulfite vs Hydrochloric acid Copper vs ...

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Sweet 16 Chemistry Ion Tournament ... The winner of the Chemistry Ion Tournament is the insoluble product obtained when the two finalists are dissolved in water. To save you time, the answer sheet is also attached. Since the presentation of solubility rules varies slightly, we strongly recommend using the solubility rules found in the

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