

The Nature Of Gases Section Review Answers

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The Nature Of Gases Section

The theory assumes that the volume occupied by a gas is mostly _____, that the particles of gas are relatively _____, move _____ of each other, and are in constant _____ motion.

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Nature Of Gases Section Review Answers

The Nature of Gases 13.1. because during an evaporation, particles with highest kinetic energy tends to escape first and the particles left have lower kinetic energy making it cool.

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Chapter 13- The States of Matter 13.1- The Nature of Gases section 13.2 the nature of liquids (pages 390-395) This section describes a model for liquids in terms of kinetic energy and the attractive forces between the particles in a liquid.

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Behavior of Gases - Ch. 8. Using the definition of pressure, explain why snowshoes prevent you from falling through the snow whereas normal snow boots do not.

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Name Date Class. THE NATURE OF GASES Section Review. Objectives □ Describe the assumptions of the kinetic theory as it applies to gases □ Interpret gas pressure in terms of kinetic theory □ Define the relationship between Kelvin temperature and average kinetic energy.

THE NATURE OF GASES - Quia

Chapter 13 States of Matter 137. SECTION 13.1 THE NATURE OF GASES (pages 385-389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385-386)

Name Date Class STATES OF MATTER 13

a. a measure of the pressure exerted by a gas above a liquid b. the forces between molecules c. the boiling point of a liquid at a pressure of 101.3 kPa d. a fluid with a fixed volume e. the change of a gas or vapor directly to a liquid Column A vapor pressure condensation normal boiling point liquid intermolecular attractions ____ 17. ____ 18. ____ 19.

05 CTR ch13 7/12/04 8:12 AM Page 317 THE NATURE OF GASES 13

Chapter 13.2 States of Matter, The nature of liquids. The particles in a liquid have an attraction for each other. The kinetic energy in them cannot overcome this attraction ... Vapor pressure is the pressure exerted by a gas above a liquid in a sealed container. More heat = more vapor pressure. Past boiling point is more vapor pressure.

Chapter 13.2 States of Matter, The nature of liquids

131 the nature of gases section review answer key FB7DF1E49821A7B8AF656B3E406CC7F5 131 The Nature Of Gases Land for food and bioenergy in the IMAGE scenario for 1.5 °C.

131 The Nature Of Gases Section Review Answer Key

gases are similar, and one way in which they are different. Both liquids and gases can flow, so they can take the shape of their container. The molecules in a liquid have intermolecular attractions that are not present in gases. Therefore, liquids have a definite volume and will not simply fill their container.

Chapter 13

Chapter 13 States of Matter137 SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance.

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