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320 Chapter 10 • The Mole Section 110.10.1 Measuring Matter MAIN Idea Chemists use the mole to count atoms, molecules, ions, and formula units. Real-World Reading Link Has your class ever had a contest to guess how many pennies or jelly beans were in a jar? You might have noticed that the smaller the object is, the harder it is to count.

### **Chapter 10: The Mole**

Stoichiometry Study Guide KEY Chemistry RHS – Mr. Moss. 1. Define the following: a. Stoichiometry- the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction. b. Mole –The SI unit used to measure the amount of a substance that contains  $6.02 \times 10^{23}$  atoms of that substance.

### **Stoichiometry Study Guide KEY Chemistry RHS Mr. Moss**

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Step 2 Determine the number of moles of the known substance, using mass-to-mole conversion.  
Step 3 Determine the number of moles of the unknown substance, using mole-to-mole conversion.  
Step 4 Determine the mass of the unknown substance, using mole-to-mass conversion.

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