

## *Teaching Transparency Formation Of Ions Answer Guide*

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### Teaching Transparency Formation Of Ions

Teaching Transparency 22 – Formation of Ions 1. calcium and oxygen 2. Yes; each contains equal numbers of protons and electrons. 3. ionization energy 4. positive; cation 5. negative; anion 6. No; it is not a stable octet of electrons. 7. No; it is not a stable octet of electrons. 8. Yes; it is a stable octet of electrons. 9.

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TEACHING TRANSPARENCY 22 Use with Chapter 7, Section 7.1 Formation of Ions 1. What are the names of the two elements shown? 2. Are the elements shown on the left sides of the two equations neutral? How can you tell? 3. What is the name for the energy needed to remove electrons from an atom, such as the Ca atom shown? 4.

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name for an ion with that kind of charge? 5. What kind of charge does the O atom take on as a result of the reaction? What is the name for an ion with that kind of charge? 6. Is the outer electron configuration of the Ca atom before the reaction a very stable one? How can you tell? 7.

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Name Date Class TEACHING TRANSPARENCY. Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_ TEACHING TRANSPARENCY 22 Chemistry: Matter and Change Teaching Transparency Worksheet 2 Use with Chapter 7, Formation of Ions Section 7.1 1. What are the names of the two elements shown?

### Chemistry Matter And Change Chapter 8 Teaching ...

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TEACHING TRANSPARENCY 24 Use with Chapter Section 7.3 Formulas for Ionic Compounds 1. How many valence electrons are there in an atom of sodium (Na)? What would a sodium atom tend to do in bonding with another atom to form an ionic compound? Why? 2. How many valence electrons are there in an atom of phosphorus (P)?

### Name TEA HIN Date Class TRANSPARENCY 24 Use with Chapter 7 ...

TEACHING TRANSPARENCY WORKSHEET The Activity Series Use with Chapter 9, Section 9.2 1. For each of the following pairs of elements, underline the one that would replace the other element in a compound. a. calcium, tin e. iron, copper b. bromine, fluorine f. iodine, chlorine c. aluminum, potassium g. silver, lead d. zinc, sodium 2.

### TEACHING TRANSPARENCY WORKSHEET The Activity Series

TEACHING TRANSPARENCY WORKSHEET Ionic Bonds Use with Chapter 7, Section 7.2 1. How many

valence electrons does a neutral magnesium (Mg) atom have? 2. What is the charge on a magnesium ion? What does magnesium have to do to form such an ion, and why does it tend to do so? 3. How many valence electrons does a single neutral chlorine (Cl) atom have? 4.

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Teaching Transparency Worksheets Chemistry: Matter and Change • Chapter 9 7 Examine the parts of the chemical equation at the top of the transparency. Use this information to answer the following questions about Equation 1 and Equation 2. 1. Write Equation 1 as a sentence. 2. Write Equation 2 as a sentence. 3. What substances are reactants in ...

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In general, ions of non metal atoms are larger than the atoms that they were formed from because non metals gain electrons when they form ions. The added electrons create greater repulsion, causing all of the electrons to move farther apart, increasing the size of the ion over the atom.

**1. 2. Li: K: O: S: 3. 4. Na: Cl**

Ion. An ion is an atom or group of atoms having an overall electrical charge. All metals form ions: they usually form positive ions. Some non-metals form ions - most of these are negative ions. Ionic Bond. It is formed by transfer of one or more electrons from one atom (or group of atoms) to another. Ionic Bond

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Ions & Their Charges Worksheet ... 13. What type of electrons, "s," "p," "d," or "f," are involved in ion formation of group IA and group IIA elements? Answer "s" 14. What does carbon, C #6, group IVA, form either +4 or -4 ions? Answer both 15. Why do elements in group VIIIA, noble gases, have zero, 0, as the charge that

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