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NCERT Solutions For Class 12 Chemistry Chapter 5 Surface Chemistry. Ans: Powdered substances have greater surface area as compared to their crystalline forms. Greater the surface area, greater is the adsorption. 5.4. In Haber's process, hydrogen is obtained by reacting methane with steam in presence of NiO as catalyst.

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NCERT Solutions Class 12 Chemistry Surface Chemistry. Here, the concentration of the adsorbate on the surface of the adsorbent increases. The concentration of the substance remains at the surface only in the case of adsorption. It does not penetrate through the surface to the bulk of the solid or liquid.

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Chemistry Class 12 NCERT Solutions: Chapter 5 Surface ...

NCERT Solutions class 12 Chemistry Surface Chemistry part 1. Part-I. 1. Distinguish between the meaning of the terms adsorption and absorption. Give one example of each. Ans. Adsorption is a surface phenomenon of accumulation of molecules of a substance at the surface rather than in the bulk of a solid or liquid.

NCERT Solutions class 12 Chemistry Surface Chemistry part 1

www.ncerthelp.com (Visit for all ncert solutions in text and videos, CBSE syllabus, note and many more) Chemistry Notes for class 12 Chapter 5 Surface Chemistry Surface Chemistry is the branch of chemistry which deals with the phenomenon that occurs on the surfaces or interfaces, such phenomenon includes corrosion. catalysis, crystallisation, etc

Chemistry Notes for class 12 Chapter 5 Surface Chemistry

NCERT Class XII Chemistry: Chapter 5 – Surface Chemistry. National Council of Educational Research and Training (NCERT) Book for Class XII Subject: Chemistry Chapter: Chapter 5 – Surface Chemistry. After studying this Unit, you will be able to. describe interfacial phenomenon and its significance;

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Class 12 Chemistry Revision Notes for Chapter 5 - Surface ...

Surface chemistry deals with phenomena that occur at the surfaces or interfaces. The interface or surface is represented by separating the bulk phases by a hyphen or a slash. For example, the interface between a solid and a gas may be represented by solid-gas or solid/gas. Due to complete miscibility, there is no interface between the gases.

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