

The Ultimate Chemical Equations H Answer Key

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The Ultimate Chemical Equations H

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The Ultimate Chemical Equations Handbook—Student Edition

The Ultimate Chemical Equations H Answer Key Ultimate equation answers max study, round 5 exercise 9—2: predict and balance the following metathesis reactions use the abbreviations (s), (l), (g), and (aq) for the

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Re-typed from The Ultimate Chemical Equations Handbook by Hague and Smith 7. Free elements, no matter how complex the molecule, have an oxidation number (valence or charge) equal to zero. OJ. Fe. T~ He < /1 1 ~ ~ L-td. 8. The following are diatomic or polyatomic elements in nature which must be committed to memory.

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The Ultimate Chemical Equations Greek, or German names. Ag Sn w (ferrum) (natrium) (argentum) (stannum) (plumbum) (wolfram) iron sodium silver tin lead tungsten (cuprum) (kalium) (hydrargyrum) (stibium) (aurium) copper potassium mercury antimony gold Symbols are used as a sort of shorthand in writing the names of elements.

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9. 10. indicate the phase of reactants and products where possible [i.e., (aq) (s) (l) (g)]. A sample of calcium carbonate is heated. $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$ Sulfur dioxide gas is bubbled through water. $\text{SO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{SO}_3(\text{aq})$ Solid potassium oxide is added to a container of carbon dioxide gas.

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Chemical Equations. First, because atoms like to have full valence shells, single H or O atoms are rare. In nature, both hydrogen and oxygen are found as diatomic molecules, H_2 and O_2 , respectively (in forming diatomic molecules the atoms share electrons and complete their valence

shells). Hydrogen gas, therefore,...

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Write a balanced molecular equation describing each of the following chemical reactions. (a) Solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas. (b) Gaseous butane, C_4H_{10} , reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor.

4.1Writing and Balancing Chemical Equations

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