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Chapter 13: Gases. 13-1: Model to Explain Gas Laws 13-2: Measuring Gas Laws 13-3: The Gas Laws 13-4: Ideal Gas Law 13-5: How Gases Work. STUDY. PLAY. volume. space a gas takes up (inversely proportional to pressure) pressure. force per area placed on an object (inversely proportional to volume) pressure.

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____ 5. alloy e. a solution containing two or more metals ____ 6. aqueous solution f. homogeneous mixture of two or more substances in a single. physical state ____ 7. electrolyte g. substance that does the dissolving in a solution. In the space provided, describe and give an example of each of the following types of solutions. 8.

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10—2 Review and Reinforcement (continued) 9. $0.0621 \text{ mol SS} = (Q \text{ } \$ \text{ gK2S } 10. 65.8 \text{gMg} = \text{l. atoms Mg}$ Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in 0.40 mol of sulfur- 12. Calculate the number of atoms in 2.30 mol of silver. G.oux 13. You have a tank of NO₂ gas.

10-2 Review and Reinforcements

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2 Answer.Key 2-1 Explore, page 5 The original source of energy for this activity ... 2-1 Review and Reinforcement, pages 8-9 . 1. energy 2. radiant 3. chemical potential 4. joule 5. created, destroyed ... 13. crystallization 14. distillation 15. chromatography 16. electrolysis 17. distillation

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5. 6. Stoichiometry is the study of the reactions. Stoichiometry can be used to determine how much product will form from a given amount of The in a balanced equation indicate(s) the number of particles of each substance taking place in the reaction. It is possible to interpret the coefficients in a balanced chemical equation as either the number

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