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$(\text{CH}_4\text{N}) \times 2 = \text{C}_2\text{H}_8\text{N}_2$  4.63 mol O \_\_\_\_ The molecular formula of a compound is either the same as its experimentally determined empirical formula, or it is a simple whole-number multiply of its empirical formula. For Carbon Dioxide, the empirical and molecular formulas are the

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Calculate the empirical formula of a compound that has a percent composition of 36.1%Ca and 63.9% Cl.  $\text{CH}_2\text{O}$  Calculate the empirical formula of a compound that has a percent composition of 3.7%H, 44.4%C, and 51.9%N? Chapter 10.3 Percent Composition and Chemical Formulas What is the

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Section 10.3 Percent Composition and Chemical Formulas 305 10.3 Percent Composition and Chemical Formulas Is your shirt made of 100 percent cotton or wool, or is the fabric a combination of two or more

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103 Percent Composition And Chemical Formulas Answer Key 10.3 percent composition and chemical formulas a molecular formula of a compound is a whole-number multiple of its empirical formula. lesson summary percent composition of a compound percent composition is the percent by mass of each element in a compound. to find

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The percent composition of the compound is:  $\text{H} = (2\text{g}/18\text{g}) \times 100 = 11.1\%$   $\text{O} = (16\text{g}/18\text{g}) \times 100 = 88.9\%$  Percent composition can be calculated the chemical formula of a compound, or it can be determined experimentally.

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