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Use VSEPR theory to predict the electron-pair geometry and the molecular geometry of iodine trichloride, ICl₃. a.The e--pair geometry is trigonal-planar, the molecular geometry is trigonal-planar. b.The e--pair geometry is tetrahedral, the molecular geometry is trigonal-pyramidal. c.The e--pair geometry is tetrahedral, the molecular geometry is trigonal-planar.

VSEPR theory? | Yahoo Answers

Best Answer: For VSEPR theory you have to account for the total number of bonds and lone pairs that surround an atom in a molecule. H₂O has an electronic geometry of tetrahedral, its molecular geometry is bent. In contrast, NH₃ has one lone pair it's electronic is also tetrahedral but its molecular geometry is trigonal pyramidal.

VSEPR theory? | Yahoo Answers

Best Answer: Valence Shell Electron Pair Repulsion Theory.

What is VSEPR theory? | Yahoo Answers

According to VSEPR theory, in which fashion will the bonds and lone pairs of electrons be arranged about the central atom in the following molecules or molecular ions? linear, octahedral, tetrahedral, trigonal bipyramidal, or trigonal planar TeF₄ molecule (The central atom is Te.) I³⁻ ion (The central atom is I.) SO₃ molecule (The central atom is S.)

VSEPR theory? | Yahoo Answers

Best Answer: The application of VSEPR theory can be expanded to complicated molecules. The application of VSEPR permits the reliable prediction of molecular structure.

VSEPR theory? | Yahoo Answers

Best Answer: well the vsepr theory just says that the electrons will want to be as far away from each other as they can get. but anyway, to get the shape of PCl₃, you first have to draw the lewis structure. to do that, you have to get the total number of valence electrons (electrons in the atom's most outer electron shell)

Use VSEPR theory to predict the ... - nz.answers.yahoo.com

Best Answer: You don't need VSEPR, Euclid pointed out that two points determine a line.

Use VSEPR theory to predict the ... - uk.answers.yahoo.com

Best Answer: HCl H-Cl (you can only arrange them in this way) CO₂ O=C=O (there are two atoms bonded to a central atom and no lone pairs)

Using the VSEPR theory, which of ... - nz.answers.yahoo.com

Best Answer: Being closer to N means that F is closer to the N lone-pair. Since the presence of this lone pair is what dictates how short the bond angle is, you can deduce that the lone pairs on F and the lone pair on N will repel more strongly than the sigma bond of H and the lone pair of N, hence the shorter bond angle.

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