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138 Chapter 5 Electrons in Atoms Electron Configurations for Elements in Period Three Table 5-4 Figure 5-19. This sublevel diagram shows the order in which the orbitals are usually filled. The proper sequence for the first seven orbitals is 1s, 2s, 2p, 3s, 3p, 4s, and 3d.

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5.1 Light and Quantized Energy. MAIN Idea Light, a form of electromagnetic radiation, has characteristics of both a wave and a particle. 5.2 Quantum Theory and the Atom. MAIN Idea Wavelike properties of electrons help relate atomic emission spectra, energy states of atoms, and atomic orbitals.

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5. In general, the higher the electron is on the energy ladder, the _____ it is from the nucleus. Scientist Model of Atom Dalton Thomson Rutherford Bohr The atom is a solid indivisible mass. The atom is a ball of positive charge with electrons stuck into the ball. Electrons are arranged in concentric circular paths around the nucleus.

SECTION 5.1 MODELS OF THE ATOM (pages 127-132)

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Chapter 5 Electrons in Atoms . Name Date 11. The number of sublevels in an energy level is equal to the square of the principal quantum number of that energy level. 12. The maximum number of electrons that can occupy the fourth principal energy level of an atom is 32. 13. The higher the energy level occupied by an electron the more

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