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### **Tonicity Calculations: An Aid to Understanding | Insight ...**

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For a solution, Raoult's law relates the relative concentrations of the components in solution with their relative vapor pressures above the solution.

### **Definition of Raoult's Law | Chegg.com**

We care about molality because freezing point depression is a colligative property, a property that depends on how many solute particles are in the solvent, not the kind of solute particles. Molality,  $m$ , is one piece of this "how many solute particles are present?" question. The Van 't Hoff factor is the second part of the "how many solute particles are present?"

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### **Definition of Acids And Bases | Chegg.com**

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