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the orientation and the energy of the colliding particles., for the reaction to occur the particles must be oriented in a good position that lets the bonds to break and the atoms to rearrange otherwise the colliding molecules would just bounce off each other.

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We can take as an example the reaction of a molecule of thioacetamide, CH3 —CS—NH2, with water to yield acetamide, CH3 CO—NH 2 and (Figure 22-1). In the original thioacetamide molecule, the central carbon atom is bound to C and to N by bonds, and by a 5, 7T double bond to S (Figure 22-1 a).

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2 Answer.Key 2-1 Explore, page 5 ... 2-1 Review and Reinforcement, pages 8-9 . 1. energy 2. radiant 3. chemical potential 4. joule 5. created, destroyed ... 22. 60.7 g 23. First remove the paper clips with a magnet or your hand. Then filter the sand from the water. Finall)" distill the salt water and collect

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16—3 Review.nd Reinforcement, pages 20-2J 5. endothermic 6. true 7. increases 8. true 9. The Concentrations of reactants and products, pressure, and temperatuÈe affect the equilibrium 10. Le Chatelier's principle states that if a change is imposed on a system at equilibrium, the equilibrium position will shift in the

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