

## *Worksheet Introduction To Specific Heat Capacities Answers*

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**Worksheet Introduction To Specific Heat**

Water has the highest specific heat capacity and metal has the lowest. 6. Here are the heat capacities of the four substances: 0.10 cal/g °C, 0.25 cal/g °C, 1.0 cal/g °C, & 0.2 cal/g °C. Match & then label each substance with its specific heat capacity on the graph. See graph above.

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Worksheet- Introduction to Specific Heat Capacities per, Heating substances in the sun: The following table shows the temperature after 10.0 g of 4 different substances have been in direct sunlight for up to 60 minutes. Sand (0 C) 400 C 450C Metal 250 35 45 55 65 c Time (minutes) 0 (initial) 15.0 min

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**Specific Heat Capacity Worksheets - Lesson Worksheets**

Worksheet- Calculations involving Specific Heat. 1. For  $q = m \cdot c \cdot \Delta T$  : identify each variables by name & the units associated with it.  $q$  = amount of heat (J)  $m$  = mass (grams)  $c$  = specific heat (J/g°C)  $\Delta T$  = change in temperature (°C) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

**Worksheet- Calculations Involving Specific Heat**

CHEM1612 Worksheet 1: Introduction to Thermodynamics Model 1: Calorimetry. Heat is not the same thing as temperature, even though in common usage these concepts are often used interchangeable. Heat is the energy transferred from one object to another due to a difference in their temperature. Heat, therefore, has units of energy (joules, J).

### **CHEM1612 Worksheet 1: Introduction to Thermodynamics Model ...**

This video will detail which parts of the worksheet to do for classwork, and provides some examples to help you. ... Worksheet - Introduction to Specific Heat Capacities K. Morris. Loading ...

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require given that the specific heat of PCI 3 is 0.874 J/g oC? 10. A quantity of water is heated from 25.0 oC to 36.4 oC by absorbing 325 J of heat energy. What is the mass of the water? 11. A 500. g sample of an unknown metal releases  $6.4 \times 10^2$  J as it cools from 55.0 oC to 25.0 oC. What is the specific heat of the sample? 12.

### **13-05,06 Heat and Heat Calculations wkst**

Because it has a lower capacity for heat, so when it gets heat, it can't hold it without changing temperature. We finish with the equation labeled like this: I then relate the high specific heat capacity of water to how it is "cooler by the lake (Michigan)" during the summer at the beach.

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