

## *7 Chemical Quantities Practice Problems Answers*

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**7 Chemical Quantities Practice Problems**

Chemistry--Unit 3: Chemical Quantities Practice Problems I. The Mole: A Measurement of Matter 1) What is the gram molecular mass of sucrose (C<sub>12</sub> H<sub>22</sub> O<sub>11</sub>)? (342.34 g/mol) 2) Calculate the gram formula mass of KMnO<sub>4</sub> and Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>. (KMnO<sub>4</sub> = 158.04 g/mol, Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> = 310.18 g/mol) 3) How many moles is  $3.52 \times 10^{24}$  molecules of water ...

**Chemistry--Chapter 7: Chemical Quantities**

Ch. 7 - Chemical Quantities - Review/Practice Problems Part I - Calculate the molar mass for each of the following compounds. Show your work with units for #3. Molar masses should be reported to the hundredths place. Include units in the answer for molar mass. 1) NaCl 2) Fe<sub>2</sub>O<sub>3</sub> 3) Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> 4) KC<sub>4</sub>H<sub>5</sub>O<sub>6</sub> 5) Mg<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> 6) SnF<sub>2</sub>

**Ch. 7 - Chemical Quantities - Review/Practice Problems**

7.3 PERCENT COMPOSITION AND CHEMICAL FORMULAS 1. A compound analyzed in a chemistry laboratory consists of 5.34 g of carbon, 0.42 g of hydrogen, and 47.08 g of chlorine. What is the percent composition of this compound? 2. Find the percent composition of a compound containing tin and chlorine if 18.35 g of the compound contains 5.74 g of tin.

**7 Chemical Quantities Practice Problems - LPS**

7 - Chemical Reactions and Quantities - Practice Problems I'm trying a different set up for the practice problems. This still contains the practice problems you need to master for the test. I've organized the problems by sections in the chapter and provided a summary that may prove useful.

**Chemical Reactions and Quantities Practice Problems**

CHEMISTRY NOTES - Chapter 7 Chemical Quantities Goals : To gain an understanding of : 1. Problem solving in chemistry. 2. The use of dimensional analysis to solve problems. 3. The concept of the mole. 4. The relationship between masses of substances and moles of substances. 5. The relationship between moles and the volumes and densities of ...

**CHEMISTRY NOTES - Chapter 7 Chemical Quantities**

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Chemical Quantities: The Mole 6 Chapter 7 Assignment & Problem Set Gas Density 27. Honors What is the density of krypton gas at STP? 28. Honors What is the density of NO<sub>2</sub> at STP? 29. Honors The density of a certain gas is 1.96 g/L. What is the molar mass of this gas?

**Chapter 7 Homework - Maine-Endwell Middle School**

Chemical Quantities - Notes Dimensional Analysis - Problem solving technique which uses conversion factors. Conversion Factor - (Equivalents) a numerical factor used to multiply or divide a quantity when converting from one system of units to. Ex., 1 foot = 12 inches Mental Note Cards - It helps to make mental note cards when solving these problems.

**Unit 7 Worksheet Answers - Chemical Quantities Notes ...**

Practice Problems. A compound formed when 9.03g Mg combines completely with 3.48g N. What is the percent composition of this compound? When a 14.2g sample of mercury (II) oxide is decomposed into its elements by heating, 13.2g of Hg is obtained. ... Chapter 10 - Chemical Quantities Last modified by:

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assign problems 1-12, 15-16. It is the mass of 1 mol of that compound. 80.1 48.0 The mass of a mole of a compound is determined by adding the atomic masses of the atoms making up the molecule. The atomic mass of an element is the mass of a single atom in atomic mass units. 0132525887\_CHEM\_WKBK\_CH 10.indd 131 3/24/10 7:36:56 PM

### Chemical Quantities - Weebly

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Chemistry--Unit 3: Chemical Quantities Practice Problems I. The Mole: A Measurement of Matter 1) What is the gram molecular mass of sucrose (C<sub>12</sub> H<sub>22</sub> O<sub>11</sub>)? 2) Calculate the gram formula mass of KMnO<sub>4</sub> and Ca<sub>3</sub> (PO<sub>4</sub>)<sub>2</sub>. 3) How many moles is  $3.52 \times 10^{24}$  molecules of water? 4) How many atoms of zinc are in 0.60 mol of zinc? II.

### Chemistry--Chapter 7: Chemical Quantities

Chapter 7 Chemical Quantities Adapted from notes by Stephen Cotton Section 7.1 ... Section 7.3 Percent Composition and Chemical Formulas ... Calculating Percent Composition of a Compound Like all percent problems: Part Whole  $\times$  . 5 Example

### Section 7.1 Chemical Quantities - Laurium - Keweenaw

Chapter 10 Chemical Quantities 247 Practice Problems In your notebook, solve the following problems. SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER 1. What is the molar mass of sucrose (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>)? 2. What is the molar mass of each of the following compounds? a. phosphorus pentachloride (PCl<sub>5</sub>) b. uranium hexafluoride (UF<sub>6</sub>) 3.

### Objectives Vocabulary Key Equations

A.P. Chemistry Classroom Technology Help en.wikipedia.com ... Chapters 7, 8 and 9 : Chemical Quantities, Reactions and Stoichiometry. Handouts Chapters 7, 8 and 9 Syllabus Section 7.1 and 7.2 Notes - Mole Conversions "Mole Capacitor" handout with demonstration practice problems Section 8.1 Notes - Writing and Balancing Reaction Equations ...

### Chapters 7, 8 and 9 : Chemical Quantities, Reactions and ...

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