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Review section 4.4 of the textbook, especially the Oxidation numbers rules on page 137. Chemistry is an experimental science where measurements are made. The theories and equations we see in the text are backed up by experimental results. The following example in the video below helped shape Electrochemistry. 20.1 Electrochemistry Observations

20.1 Oxidation States and Oxidation-Reduction Reactions ...

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shown in Figure 20-2. The formation of the covalent bond by sharing of elec-trons also is an oxidation-reduction reaction. 636 Chapter 20 Redox Reactions Figure 20-1 The reaction of magnesium and oxygen involves a transfer of electrons from magnesium to oxygen. Therefore, this reaction is an oxidation-reduction reac-tion. Using the ...

Chapter 20: Redox Reactions - Neshaminy School District

12—2 Review and Reinforcement (continued) 10. 20.0 g of potassium react with water to produce potassium hydroxide and hydrogen gas. How many liters would the hydrogen gas occupy at STP? 573 L u 301, 1 2 k 11. If 30.2 g of aluminum react with HCl to produce alurninum chloride and hydrogen gas, how

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3-1 review and reinforcement. STUDY. PLAY. Matter is neither created or destroyed in chemical reactions. Lavoisier. A given compound always has the same relative numbers and kind of atoms. Dalton. All atoms of a given element are identical, but they differ from those of any other element.

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Answer Key

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Name _____ Date _____ Class 10-1 Review and Reinforcement (continued) Find the formula mass of each of the following compounds. Show all your work. 11. CaCO_3

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