

## *Worksheet Calculations Involving Specific Heat Answer Key*

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### Worksheet Calculations Involving Specific Heat

Worksheet- Calculations involving Specific Heat. 1. For  $q = m \cdot c \cdot \Delta T$  : identify each variables by name & the units associated with it.  $q$  = amount of heat (J)  $m$  = mass (grams)  $c$  = specific heat (J/g°C)  $\Delta T$  = change in temperature (°C) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

### Worksheet- Calculations involving Specific Heat - Mr. Winters

Showing top 8 worksheets in the category - Specific Heat. Some of the worksheets displayed are Name per work introduction to specific heat capacities, Work calculations involving specific heat, Specific heat wksht20130116145212867, Specific heat, , Specific heat practice work, Specific heat work, Chemistrytemperaturespecificheatwork answer key.

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Showing top 8 worksheets in the category - Calculating Heat. Some of the worksheets displayed are Work calculations involving specific heat, 13 0506 heat and heat calculations wkst, 13 06abc heat and heat calculations wkst key, Name per work introduction to specific heat capacities, Calculating heat, Specific heat wksht20130116145212867, Whole house heat loss gain work, Chemistry ii enthalpy work name.

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### Name: Per: Worksheet- Introduction to Specific Heat Capacities

Worksheet - Calculations Involving Specific Heat. 1. For  $q = m \cdot c \cdot \Delta T$  : identify each variable by name & the units associated with it. 2. Heat is not the same as temperature, yet they are related. Explain how they differ. a. Perform calculations using ( $q = m \cdot c \cdot \Delta T$ ) b.

### Worksheet - Calculations Involving Specific Heat

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### Worksheet Calculations Involving Specific Heat ...

Worksheet- Calculations involving Specific Heat 1. For  $q = m \cdot c \cdot \Delta T$  : identify each variables by name & the units associated with it.  $q$  = amount of heat (J)  $m$  = mass (grams)  $c$  = specific heat (J/g°C)  $\Delta T$  = change in temperature ( °C) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

### 1025\_Lec 2\_01\_Specific Heat Worksheet with answers ...

Worksheet- Introduction to Specific Heat Capacities Heating substances in the sun: The following table shows the temperature after 10.0 g of 4 different substances have been in direct sunlight for up to 60 minutes.

### Worksheet- Introduction to Specific Heat Capacities

Two page worksheet using Specific Heat Capacity. Questions start easy then become gradually harder. Answers included on separate sheet. Also includes a spreadsheet to show how the calculations have been done.

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Name: Worksheet- Introduction to Specific Heat Capacities Per: Heating substances in the sun: The following table shows the temperature after 10.0 g of 4 different substances have been in direct sunlight for up to 60 minutes.

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Heat Transfer/ Specific Heat Problems Worksheet Solving For Heat (q) 1. How many joules of heat are required to raise the temperature of 550 g of water from 12.0 oC to 18.0 oC? 2. How much heat is lost when a 64 g piece of copper cools from 375 oC, to 26 C? (The specific heat of copper is 0.38452 J/g x oC). Place your answer in kJ. 3.

### **Heat Transfer/ Specific Heat Problems Worksheet**

Here are the heat capacities of the four substances: 4.18 J/g o c, 1.0 J/g o c, 0.80 J/g o c, & 0.60 J/g o c. March then /abe/ each substance wi its specific h t capacity on he graph. If something has a high specific heat capacity will it take a lot of heat or a little heat to change its

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require given that the specific heat of  $\text{PCl}_3$  is 0.874 J/g oC? 10. A quantity of water is heated from 25.0 oC to 36.4 oC by absorbing 325 J of heat energy. What is the mass of the water? 11. A 500. g sample of an unknown metal releases  $6.4 \times 10^2$  J as it cools from 55.0 oC to 25.0 oC. What is the specific heat of the sample? 12.

### **13-05,06 Heat and Heat Calculations wkst**

– worksheet calculations involving specific heat winters 5 copper has a specific heat of 0 385 j g $\times$ °c a piece of copper absorbs 5000 j of energy and undergoes a temperature change from 100 °c to 200 °c what is the mass of the piece of copper  $q = 5000 \text{ j}$   $m = c = 0 \text{ 385 j g}^\circ\text{c}$   $\Delta t = 200^\circ\text{c}$   $100^\circ\text{c} = 100^\circ\text{c}$   $m = 129 \text{ 87 g}$  100 g endothermic or ...

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