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10—2 Review and Reinforcement (continued) 9. $0.0621 \text{ mol SS} = (Q \$ \text{gK}2\text{S } 10. 65.8\text{gMg} = \text{l. atoms Mg}$ Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in 0.40 mol of sulfur- 12. Calculate the number of atoms in 2.30 mol of silver. G.oux 13. You have a tank of NO₂ gas.

10-2 Review and Reinforcements

11—2 Review and Reinforcement Solving Stoichiometry Problems If the statement is true, write

"true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided. 1. 2. 5. The major types of stoichiometry problems are mass-mass, mass-volume, and volume-volume.

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3—3 Review and Reinforcement Modern Atomic Theory Complete the following table. 1. 2. 3. If the statement is true, write "true." If it is false, change the underlined word or words to make the statement true. Write your answer on the line provided. 4. Moseley discovered that all atoms of an element have the same number of neutrons in ...

Name Date Class 3—3 Review and Reinforcement

7-2 Review and Reinforcement. Covalent Bonding. ... Answer the following questions as directed. 21. Explain why the molecules SF₄ is an exception to the octet rule.

2 3 Review And Reinforcement Answer Key

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Write your answer on the line provided s. 10. 12. The term stoichiometry is derived from the words stoicheion, meaning element, and metron, meaning measure. You can determine the number of moles Of any substance produced in a reaction if you know the number Of moles of at least of the reactants.

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16—3 Review and Reinforcement, pages 20-2J 5. endothermic 6. true 7. increases 8. true 9. The Concentrations of reactants and products, pressure, and temperature affect the equilibrium 10. Le Chatelier's principle states that if a change is imposed on a system at equilibrium, the equilibrium position will shift in the

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11—1 Review and Reinforcement (continued) Solve each of the following problems as directed. Show all your work. 13. How many moles Of magnesium are required to react with 2.0 mol of hydrochloric acid? The equation for this reaction is $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ 14. Aluminum reacts with HCl to produce aluminum chloride (AlCl₃) and hydrogen gas.

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