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16 Colligative Properties Section Review

SECTION 16.3 COLLIGATIVE PROPERTIES OF SOLUTIONS (pages 487–490) This section explains why a solution has a lower vapor pressure, an elevated boiling point, and a depressed freezing point compared with the pure solvent of that solution. Vapor Pressure Lowering (pages 487–488) 1. Properties of a solution that depend only on the number of ...

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is given by $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$ Because $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is a strong electrolyte, it dissociates completely into $\text{NH}_4^+(\text{aq})$ and $\text{SO}_4^{2-}(\text{aq})$ ions. Assume a one kilogram solution. The number of moles of ions in one ...

CHAPTER 16. Colligative Properties of Solutions

Solutions 487 Print ✎ Guided Reading and Study Workbook, Section 16.3 ✎ Core Teaching Resources, Section 16.3 Review ✎ Transparencies, T175DT176 Technology ✎ Interactive Textbook with ChemASAP, Assessment 16.3 ✎ Go Online, Section 16.3 16.3 FOCUS Objectives 16.3.1 Identify three colligative prop-erties of solutions.

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Chapter 16

This section introduces a third category that is a subset of the intensive properties of a system. This

third category, known as colligative properties , can only be applied to solutions. By definition, one of the properties of a solution is a colligative property if it depends only on the ratio of the number of particles of solute and solvent ...

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Colligative properties depend only on the number of dissolved particles (that is, the concentration), not their identity. Raoult's law is concerned with the vapour pressure depression of solutions. The boiling points of solutions are always higher, and the freezing points of solutions are always lower, than those of the pure solvent.

Colligative Properties of Solutions - Introductory ...

Chapter 16 Solutions 403 Section Review Objectives • Identify the three colligative properties of solutions • Describe why the vapor pressure, freezing point, and boiling point of a solution differ from those properties of the pure solvent. Vocabulary • colligative properties • freezing-point depression • boiling-point elevation Part ACompletion Use this completion exercise to check ...

05 CTR ch16 7/12/04 8:14 AM Page 399 PROPERTIES OF ...

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