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10 2 Review And Reinforcement

10—2 Review and Reinforcement (continued) 9. $0.0621 \text{ mol SS} = (Q \text{ } \$ \text{ gK2S } 10. 65.8 \text{ gMg} = \text{I. atoms Mg}$ Solve each of the following problems as directed. Show all your work. 11. Calculate the number Of atoms in 0.40 mol of sulfur- 12. Calculate the number of atoms in 2.30 mol of silver. G.oux 13. You have a tank of NOZ gas.

10-2 Review and Reinforcements - Front Page

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Worksheet # 5 Quarter 3 Packet 2: Mole Conversions 10-2 Review and Reinforcement Packet #2 Quarter 3 assignment 5. Mole Conversions Answer Each of the following questions in the space provided. 1. How would you calculate the number of moles present in a given mass of a substance? _____ 2.

Mole Conversions - OoCities

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10 1 REVIEW AND REINFORCEMENT CHEMICAL MEASUREMENTS ...

12—2 Review and Reinforcement (continued) 10. 20.0 g of potassium react with water to produce potassium hydroxide and hydrogen gas. How many liters would the hydrogen gas occupy at STP? 573 L u 301, 1 2 k 11. If 30.2 g of aluminum react with HCl to produce aluminum chloride and hydrogen gas, how

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2 Answer.Key 2-1 Explore, page 5 The original source of energy for this activity would be the stored chemical energy in the ... 2-1 Review and Reinforcement, pages 8-9 . 1. energy 2. radiant 3.

chemical potential 4. joule 5. created, destroyed 6. natural gas 7. kinetic . 8. f . 9. c 10. b 11. a . 12. e .

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