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1B)

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[2.] Write sho a

NS: y FIVE (5) questions:

LQ ace ia Ala atomic nurnbur and mass number.

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il What is the difference between physical and chemical properties? Give an example of each.

(iit) Why ionic compounds have high meting «nd boiling points?

Write down electronic configuration of akiminium according to shells and subshells.

iv) Differentiate between shell and subsneel with examples.

Wi Dexcribe the trend of electronegativity

(vil) Define ionization energy.

(vill) Write two features of long form of periodic fee 8 com

SY Write short answers to Ni Cy

(1) Define octet

sq

Te Valent bond oe one example.

i NN at is hydrogen bonding?

liv} What is absolute zero?

How intermolecular forces affects evaporation?

(vi) Define aqueous solution with ane exanipl:

What is percentage - mass/volume (% m/v)? .

(Vill) Define saturated and unsaturated solutior .

[4] Write short answers to any FIVE «5) questions:

(4) Define reducing agent.

(ii) What is meant by rust?

(it) Define electrochemical cell.

(iv) What is meant by HS AND le.

= vt is Copper sien Glectacal wires?

ve at is Son by electropositivity?

i Write down the reaction of fluorine with w.ll@f,

PART ~ II

Note: Attempt any Two questions.

(5.)(a) What were the results concluded by Rutherford in his experiment? ®

(b) Define molecule and explain its types. 5

(6.)(a) Explain dipole-dipole interaction with example of HC (.
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(b) State Boyle's law and derive its equation.

[7.](a) Define electroplating and explain the electroplating cell.

(b) What is molarity? write its unit or one molar

mass of solute (OH) 3 a

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