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# Hydrogen adsorption and dissociation on $\text{Au}_n\text{Y}$ ( $n = 1-12$ ) nanoclusters: a DFT investigation†

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The interaction between nanomaterial systems and hydrogen has recently emerged as a compelling study model, offering valuable insights for designing materials with applications in nanotechnology, catalysis, and energy storage. Among these, transition metal-doped gold clusters exhibits intriguing stability and electronic properties, making them promising candidates for hydrogen related processes. In this study, we employ density functional theory (DFT) to investigate the interaction between  $\text{H}_2$  molecules and small-sized gold clusters doped with yttrium. Our finding reveals that most of the bare cluster structures remain intact during the  $\text{H}_2$  adsorption, regardless of whether the process occurs molecularly or dissociatively. A comprehensive analysis indicates that the preferred adsorption configuration is governed by multiple factors, including adsorption site (surface vs. encapsulated), relative electronegativity, and the atomic coordination number. The calculations demonstrate that dissociative adsorption of  $\text{H}_2$  on  $\text{Au}_6\text{Y}$  and  $\text{Au}_{11}\text{Y}$  clusters is both thermodynamically and kinetically favorable. However, for  $\text{Au}_n\text{Y}$  ( $n = 1, 4-5, 8, 10$ ) dissociative adsorption is hindered by a significant energy barrier before reaching the final state, while for species with  $n = 2, 3, 7$ , and 9 molecular adsorption is more favorable due to intrinsic energy preferences. This study provides fundamental insights into the adsorption sites and detailed analysis of adsorption kinetics on  $\text{Au}_n\text{Y}$  clusters, laying the groundwork for further theoretical and experimental investigations into the hydrogenation process in nanostructured materials.

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## 1. Introduction

Nowadays, hydrogen has been envisioned as a key energy carrier for sustainable development of the future economy, meeting essential environmental and energy requirements.<sup>1-3</sup> Compared to traditional fuels like gasoline ( $44 \text{ MJ kg}^{-1}$ ) and diesel ( $45.6 \text{ MJ kg}^{-1}$ ),<sup>4,5</sup> hydrogen offers the highest energy density per unit mass ( $\sim 120 \text{ MJ kg}^{-1}$ ) and conversion efficiency and can be produced through water-splitting reactions without  $\text{CO}_2$  emissions. Moreover, it can be conveniently stored in various forms, including gaseous, liquid, or metal hydrides.<sup>6-8</sup> However, the practical implementation of hydrogen energy storage after generation and during transport faces significant challenges, primarily due to its kinetic stability under ambient conditions

and the low energy density per unit volume of solid storage materials.

To address this bottleneck, various strategies have been explored.<sup>9-12</sup> Among them, metal hydride materials have garnered special attention as model solid-state hydrogen storage systems, offering a more compact and secure alternative. They provide the possibility of preventing critical operating conditions like high pressures and elevated temperatures, which are typical of conventional compressed gas and liquefied hydrogen storage.<sup>13</sup> To enhance the practical prospect of metal hydride, a deeper understanding of the interaction between hydrogen and potential surface at the molecule level is essential. Therefore, nanoclusters consisting of a few atoms have emerged as promising candidates for providing fundamental insights into potentially reactive sites, given significant variation in physical and chemical properties with size and composition.

Among various metal nanoclusters, gold nanoclusters ( $\text{Au}_n$ ) are considered particularly advantageous due to their superior stability, abundance, and cost-effectiveness, even surpassing palladium and platinum in many aspects. While gold is chemically inert in its bulk form, it exhibits remarkable catalytic activity at the nano scale. For example, gold nanoclusters have shown promising performance in catalytic processes such as water-splitting reaction<sup>14</sup> and CO oxidation.<sup>15</sup> Interestingly, this

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catalytic behavior is highly dependent on their geometric structure. Bootharaju and co-workers<sup>16</sup> investigated the  $\text{Au}_{12}\text{Ag}_{32}/\text{TiO}_2$  system, which features a gold icosahedral core encapsulated by a silver dodecahedral shell. Their finding revealed a substantial enhancement in solar  $\text{H}_2$  production efficiency, approximately 6.2 and 37.8 times higher than those of  $\text{Ag}_{44}/\text{TiO}_2$  and bare  $\text{TiO}_2$ , respectively. Furthermore, key factors influencing the reactivity and catalytic potential of gold nanoclusters include cluster size, atomic composition, and charge state.<sup>17–19</sup> By tailoring these parameters, such as through doping or precise control of cluster size, researchers can design tunable materials optimized for specific catalytic processes. For instance, both experimental and theoretical studies have explored  $\text{O}_2$  activation on small-sized anionic  $\text{Au}_n^-$  clusters.<sup>20</sup> S. Pal and D. Manzoor have compared the catalytic efficiency of  $\text{Au}_8$ ,  $\text{Pd}_8$ ,  $\text{Au}_{8-n}\text{Pd}_n$  ( $n = 1-7$ ),  $\text{Au}_7\text{Si}$ ,<sup>21,22</sup>  $\text{Au}_{18}$  and endohedral  $\text{Au}_{18}\text{M}$  ( $\text{M} = \text{Na}, \text{K}, \text{Mg}, \text{Ca}, \text{Al}, \text{Ga}$ ) cages,<sup>23</sup> and Al-, Hf-, and Ge-doped  $\text{Au}_{20}$  cages.<sup>24</sup> Similarly, Zeng and co-workers<sup>25</sup> using DFT calculation studied the catalytic behavior of sub-nanometer-sized gold nanoclusters supported on  $\text{TiO}_2$ . They observed that cluster shape and size significantly influenced catalytic performance, with adsorption energies for  $\text{O}_2$  and CO decreasing as the cluster size increased within the range of 16–35 atoms.<sup>15</sup> Recent studies have also highlighted the emerging role of hydrogen atoms as dopants in gold nanoclusters.<sup>26</sup> Beyond their electronegativity, hydrogen atoms can induce chemisorption and activation on previously inert or closed-shell gold clusters,<sup>27</sup> modulated by cluster shape,<sup>28,29</sup> and charge state.<sup>30</sup> This, in turn, enhances the adsorption and activation of  $\text{O}_2$  and CO molecules, opening new directions for designing advanced gold-based catalysts.

Research efforts have expanded beyond the studies of  $\text{O}_2$  adsorption and CO oxidation to include investigations of metal nanoclusters and their alloys in order to deepen our understanding of the kinetics and dynamics of hydrogenation reactions. Numerous density functional theory (DFT)<sup>31–34</sup> studies have investigated the adsorption and dissociation of hydrogen on metal nanoclusters. The  $\text{H}_2$  absorption process has been found to strongly depend on the physical and chemical properties, structural, size, and composition of nanoclusters, as demonstrated in 3d transition metal (Sc–Zn)-doped Cu clusters<sup>35</sup> and  $\text{Au}_{24}\text{M}$ ,  $\text{Au}_{36}\text{M}$  ( $\text{M} = \text{Pt}$  and  $\text{Pd}$ ).<sup>36</sup> Theoretical studies have demonstrated that the incorporation of a Pt dopant can significantly alter the electronic structure of gold nanoclusters  $\text{Au}_{n+1}$  ( $n = 1-12$ ), thereby enhancing their reactivity toward  $\text{H}_2$  molecules.<sup>37</sup> Li and coworkers investigated the effect of adsorbed hydrogen atoms on the geometric and electronic structures of alkali-doped gold nanoclusters, providing insights into hydrogen-cluster interaction.<sup>38</sup> Additionally, the adsorption and dissociation behavior of  $\text{TiMg}_n$  ( $n = 1-12$ ) clusters has been examined DFT, revealing key factors that govern their hydrogen activity.<sup>39</sup> Furthermore, mass spectrometry and infrared multiple photon dissociation spectroscopy combined with theoretical calculations, have provided insights into the competition between molecular and dissociative hydrogen chemisorption, which is highly dependent on the size of cationic aluminum-doped Rh clusters.<sup>40,41</sup> Recently, Lan *et al.*

conducted comparative studies on the influence of 3d transition metal atoms on hydrogen adsorption and dissociation by examining size-dependent  $\text{Ag}_n\text{Cr}$  ( $n = 2-12$ ) clusters and composition-dependent  $\text{Au}_9\text{M}^{2+}$  ( $\text{M} = \text{Sc-Ni}$ ) clusters.<sup>42,43</sup> Their research has revealed that  $\text{Ag}_3\text{Cr}$ ,  $\text{Ag}_6\text{Cr}$ , and  $\text{Au}_9\text{Ti}^{2+}$  clusters exhibit the ability to adsorb and dissociate  $\text{H}_2$ , leading to the formation of metal hydrides even at room temperature.

Additionally, 4d transition metal atoms have shown significant potential for hydrogen storage, attributed to their enhanced stability, fully filled electron configurations across multiple energy levels, diverse oxidation states, and distinct chemical properties compared to their 3d counterparts.<sup>44,45</sup> Moreover, the lower electronegativity of 4d transition metals relative to 3d elements influences their interaction with hydrogen, raising several intriguing questions, such as how charge transfer occurs in clusters and the nature of interactions between dopant clusters and hydrogen molecules. To address these questions, the effects of boron and 4d transition metal (Y–Mo) on hydrogen adsorption and storage properties have been systematically investigated.<sup>45–47</sup> Notably, the impact of charge transfer, electronegativity, and the atomic radius of the dopant elements on the catalytic mechanism and hydrogen storage capacity of  $\text{MgH}_2$  clusters has been clarified.<sup>48,49</sup> C.S. Sergio *et al.* have clarified the correlation between hydrogen storage capacity and the surface binding energy of 4d transition metal-doped carbon nanoflakes.<sup>50</sup> Recent work by J. Barabás *et al.*, combining gas-phase reaction investigations and DFT computations, has demonstrated that the catalytic activity of gold cation clusters ( $\text{Au}_n^+$ ) and yttrium-doped gold ( $\text{Au}_{n-1}\text{Y}^+$ ,  $n = 4-20$ ) clusters can be effectively tuned by cluster size, charge state, and dopant introduction.<sup>51</sup> While initial studies have explored the influence of 4d transition metals on cluster stability and reactivity, a systematic investigation of the hydrogen storage mechanism in noble metal clusters doped with 4d transition metals remains lacking. A deeper understanding of the kinetics and dynamics of hydrogenation reactions in precious metal species and their alloys could pave the way for new research directions in hydrogen storage materials.<sup>52,53</sup>

Driven by the pursuit of nanocluster materials for efficient hydrogen storage, this study aims to theoretically explore the critical role of yttrium doping in  $\text{H}_2$  adsorption and dissociation on small gold clusters. Using DFT calculations,<sup>54</sup> we systematically analyze the geometrical, relative stability, bonding characteristics, electronic properties, and adsorption energies of  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) clusters. The insights gained from this study provide a deeper understanding of the influence of 4d transition metal dopants on hydrogen adsorption and dissociation mechanisms. Furthermore, these findings offer valuable guidelines for future experimental studies, aiding the rational design of advanced nanomaterials for hydrogen storage applications.

## 2. Computational method

This work uses the DFT calculations implemented in the Gaussian 09 program<sup>55,56</sup> in combination with GaussView, to predict the geometric and electronic structures of  $\text{Au}_n\text{Y}$  ( $n = 1-$

12) clusters in their ground state, along with their molecularly and dissociatively adsorbed variants  $\text{Au}_n\text{Y-H}_2$  and  $\text{Au}_n\text{Y-2H}$ . To generate initial cluster structures, two complementary approaches were employed: (1) a stochastic algorithm was used to explore a broad range of feasible  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) configurations;<sup>57</sup> (2) manual structure construction, in which a Y atom systematically replaced an Au atom at all possible positions within the cluster. Subsequently, hydrogen adsorption configurations were investigated based on the lowest-energy  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) structures, where a single  $\text{H}_2$  molecule or two dissociated H atoms were attached at various adsorption sites.

Computationally, predicted structures were first optimized by the TPSSPTSS functionals in combination with the cc-pVDZ-pp basis set for Au and Y atoms, and LanL2DZ for H atoms. For each cluster size, various structural and spin configurations were explored and subjected to the convergence criteria. Nevertheless, only species with relative energies less than 2.0 eV were selected for single-point energy recalculations at the same functional level but with an enhanced basis set, specifically cc-pVTZ-pp for Au and Y atoms and LanL2DZ for H atoms. This approach represents a balanced trade-off between computational accuracy and efficiency and has been shown to be reliable for 4d transition metal-doped noble metal systems.<sup>58</sup> To validate the accuracy and reliability of our computational methodology, we compared binding energies (BE, eV) of  $\text{Au}_2$ ,  $\text{AuY}$ ,  $\text{AuH}$ , and  $\text{YH}$  dimers across different functionals, including TPSSPTSS, BP86, BPW91, BLYP, and PBEPBE. Tables 1 and S1† summarize the results and benchmarks them against available experimental data, confirming that the TPSSPTSS/cc-pVTZ-pp (for Au and Y) and LanL2DZ (for H) combination yields the closest agreement with experimental values. The self-consistent field (SCF)<sup>59</sup> calculations were conducted with a convergence threshold of  $2.0 \times 10^{-5}$  Hartree per Å for gradients and  $5.0 \times 10^{-3}$  Å for displacements. The stability of the lowest-energy structures was further confirmed by calculating harmonic vibrational frequencies, ensuring they correspond to true minima on the potential energy surface. To reinforce the robustness of the ground-state optimization, multiple spin multiplicities were examined for each structure. The electronic configurations of all lowest-energy structures were analyzed using the natural bond orbital (NBO) method.<sup>60</sup> To explore the reaction mechanisms, transition-state (TS) geometries were identified using the quadratic synchronous transit (QST) method,<sup>61,62</sup> with relaxed potential energy surface (PES) scans along appropriate internal coordinates. To further verify the transition states, intrinsic reaction coordinate (IRC) calculations were performed, ensuring a continuous and valid connection between reactants and products.

### 3. Optimized geometrical structures

We first optimized the lowest-energy structures of bare  $\text{Au}_n\text{Y}$  clusters and then examined both molecular and dissociative adsorption of hydrogen, leading to the formation of  $\text{Au}_n\text{Y-H}_2$  and  $\text{Au}_n\text{Y-2H}$  clusters, respectively. To ensure the stability of the obtained geometries, we conducted frequency calculations, confirming the absence of imaginary frequencies in all optimized structures. While multiple geometrical and spin isomers exist for each stoichiometry, only the lowest-energy structures of  $\text{Au}_n\text{Y}$ ,  $\text{Au}_n\text{Y-H}_2$ , and  $\text{Au}_n\text{Y-2H}$  clusters were selected for further discussion of hydrogen adsorption behavior, as shown in Fig. 1. Other low-lying isomers for  $\text{Au}_n\text{Y-H}_2$  and  $\text{Au}_n\text{Y-2H}$  clusters are provided in the ESI.†<sup>46</sup>

Our results indicate that the stable spin configurations of  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) clusters follow an alternating pattern with a singlet spin state for odd  $n$  and a doublet for even  $n$ . The lowest-energy  $\text{Au}_n\text{Y}$  structures tend to adopt a cage-like form as the cluster size increases, where the Y atom prefers a highly coordinated position. Interestingly, the structural framework of  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) clusters undergoes significant reconstruction upon yttrium doping, deviating from the planar configuration of pure  $\text{Au}_n$  clusters ( $n \leq 11$ ).<sup>64</sup> Unlike pure gold clusters, where the first 3D structure appears at  $n = 12$ , the introduction of a Y atom stabilizes 3D geometries at much smaller sizes, starting from  $n = 4$ , where four Au atoms coordinate around the Y dopant. The lowest-lying isomers of  $\text{Au}_n\text{Y}$  ( $n = 5-8$ ) retain a 2D morphology, whereas larger clusters ( $n = 9-12$ ) exhibit 3D structural preferences. These findings align well with theoretical predictions and far-infrared spectral studies of yttrium-doped gold clusters.<sup>65</sup>

Our calculations suggest that hydrogen adsorption does not significantly alter the geometric structure of bare  $\text{Au}_n\text{Y}$  clusters, except for  $\text{Au}_5\text{Y}$ , where  $\text{H}_2$  preferentially attaches to energetically unstable bridge sites (Au-Y and Au-Au). Similar behavior has been observed in transition metal-doped clusters.<sup>41-43,66</sup> For  $\text{Au}_n\text{Y}$  ( $n = 1-5, 7, 9$ , and  $12$ ), the hydrogen molecule favors direct attachment to the Y dopant, attributed to differences in electronegativity  $\chi$  between Y (1.22), H (2.2), and Au (2.54). Fundamentally, hydrogen bonding arises from electrostatic attraction, facilitated by charge transfer between atoms with significant electronegativity differences. Consequently, bonds between the low-electronegativity Y atom and high-electronegativity H atom are more favorable, in agreement with previous reports, where surface impurities are among the most favored adsorption sites for a hydrogen molecule on transition metal doped clusters.<sup>41-43,67</sup> Furthermore, hydrogen adsorption is not solely governed by electronegativity differences but also strongly influenced by atomic coordination number (N). Lower

**Table 1** Calculated binding energies (BE, eV) for AuH and YH dimers using various functionals in combination with cc-pVTZ-pp for Au and Y atoms, and LanL2DZ/SDD for H atoms

Dimer	BP86		PW91		BLYP		PBEPBE		TPSSPTSS		Experiment <sup>63</sup>
	LanL2DZ	SDD	LanL2DZ	SDD	LanL2DZ	SDD	LanL2DZ	SDD	LanL2DZ	SDD	
AuH	3.07	3.11	3.02	3.00	2.94	2.97	2.94	2.98	3.00	3.02	$3.03 \pm 0.08$
YH	2.90	3.12	2.74	2.82	2.88	3.02	2.79	2.85	2.94	3.23	—



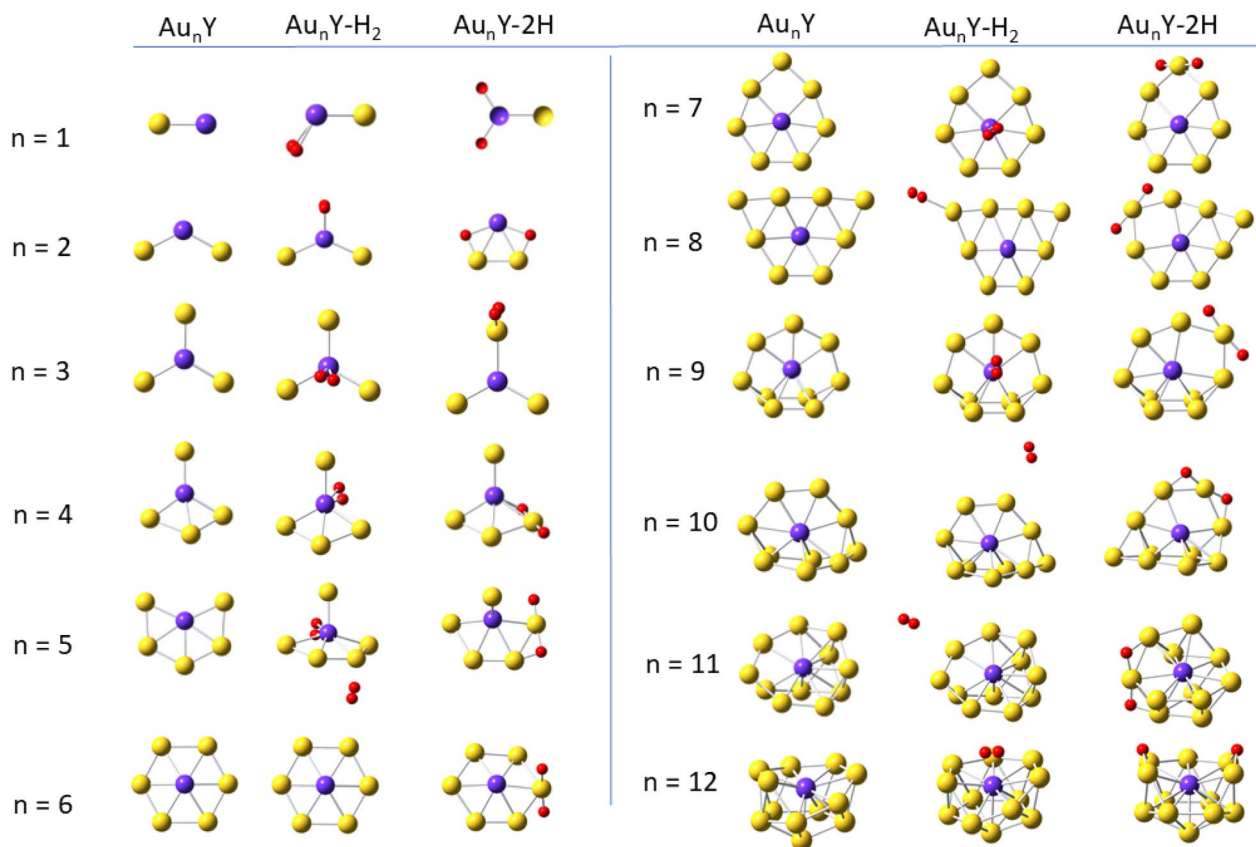


Fig. 1 Optimized structures of  $\text{Au}_n\text{Y}$ ,  $\text{Au}_n\text{Y-H}_2$ , and  $\text{Au}_n\text{Y-2H}$  clusters ( $n = 1-12$ ). The yellow, purple, and red spheres represent Au, Y, and H atoms, respectively.

coordination numbers correspond to higher surface activity, facilitating stronger  $\text{H}_2$  adsorption.<sup>68</sup> Consequently, atomic clusters with low-coordination sites exhibit enhanced hydrogen interaction capability. For example, experimental and theoretical studies on singly rhodium-doped cationic aluminum clusters ( $\text{Al}_n\text{Rh}^+$ ) using time-of-flight mass spectrometry and infrared multiple photon dissociation, in combination with DFT calculations, have demonstrated that hydrogen preferentially binds to the Al atom rather than the Rh impurity in cluster sizes where Rh is coordination-saturated or encapsulated.<sup>41</sup> This principle helps explain the behavior observed in  $\text{Au}_6\text{Y-H}_2$ ,  $\text{Au}_8\text{Y-H}_2$ ,  $\text{Au}_{10}\text{Y-H}_2$ , and  $\text{Au}_{11}\text{Y-H}_2$  clusters, where a competition between electronegativity and coordination occurs as  $\text{H}_2$  preferentially attaches to a low-coordination Au vertex ( $N_{\text{Au}} = 2$  or 3) instead of the highly coordinated Y impurity ( $N_{\text{Y}} = 6, 8$ , or 9). A similar hydrogen adsorption trend has been reported for the clusters  $\text{Al}_n\text{Cr}$ ,  $\text{Ag}_n\text{Cr}$ ,  $\text{Au}_9\text{M}^{2+}$ ,  $\text{Al}_n\text{Rh}^{2+}$ , and  $\text{Mg}_n\text{Co}$ , further reinforcing the critical role of coordination number in determining the preferred hydrogen adsorption site.<sup>41-43,66,69</sup>

Upon molecular adsorption, the hydrogen molecule may dissociate, triggering a structural transformation of the bare clusters.<sup>70</sup> However, the ground-state geometry of  $\text{Au}_n\text{Y}$  clusters generally remains intact upon hydrogen adsorption, with only minor distortions observed in  $\text{Au}_2\text{Y-2H}$ ,  $\text{Au}_5\text{Y-2H}$ , and  $\text{Au}_8\text{Y-2H}$ . As illustrated in Fig. 1, following  $\text{H}_2$  dissociation one Au atom shifts closer to another Au atom in  $\text{Au}_2\text{Y-2H}$ , while in

$\text{Au}_5\text{Y-2H}$  and  $\text{Au}_8\text{Y-2H}$ , a single Au atom is slightly displaced from two-membered and three-membered Au chains, respectively. Notably, both cluster size and the position of the Y dopant significantly influence  $\text{H}_2$  dissociation and adsorption behavior, which can be categorized into two distinct trends. For small clusters ( $n = 2-4$ ), individual hydrogen atoms preferentially bind to the surface Y atom, typically at the Au-Y bridge site. As the cluster size increases ( $n = 5-12$ ), the Y atom becomes progressively encapsulated by surrounding Au atoms, leading to H atoms favoring Au-Au bridge adsorption rather than direct interaction with Y. This trend is in good agreement with previous investigations.<sup>43</sup> In the case of  $\text{AuY-2H}$ , both H atoms preferentially adsorb on top of the Y surface site, highlighting the role of dopant coordination and atomic arrangement in governing hydrogen interaction dynamics.

## 4. Relative stability

The relative stability of a clusters is a key parameter that provides insight into its bonding or dissociation energy, which are critical in determining reactivity and stability in chemical reactions. This stability is typically assessed through average binding energies (BE, in eV) calculations. The BE is defined as the mean energy difference between the sum of the total energies of all isolated constituent atoms and the total energy of the cluster. The calculation follows the equations below:





$$\text{BE}(\text{Au}_n\text{Y}) = \frac{1}{n+1} [nE(\text{Au}) + E(\text{Y}) - E(\text{Au}_n\text{Y})] \quad (1)$$

$$\text{BE}(\text{Au}_n\text{Y} - \text{H}_2) = \frac{1}{n+3} [nE(\text{Au}) + E(\text{Y}) + 2E(\text{H}) - E(\text{Au}_n\text{Y} - \text{H}_2)] \quad (2)$$

$$\text{BE}(\text{Au}_n\text{Y} - 2\text{H}) = \frac{1}{n+3} [nE(\text{Au})_{E(\text{Y})} + 2E(\text{H}) - E(\text{Au}_n\text{Y} - 2\text{H})] \quad (3)$$

Here,  $E$  represents the total energy of the clusters or individual atoms. The binding energy (BE) values of  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) clusters range from 1.59 to 2.58 eV, as shown in Fig. 2 and Table 2. A notable increase in BE is observed for bare  $\text{Au}_n\text{Y}$  clusters in the small-size range ( $n = 1-3$ ), rising sharply from 1.59 eV to a peak of 2.55 eV. This trend persists in the presence of molecular hydrogen adsorption, although the increase in BE (from 1.96 eV to 2.48 eV) remains less pronounced compared to

bare clusters of the same size. In contrast, for dissociative adsorption ( $\text{Au}_n\text{Y}-2\text{H}$ ), BE values gradually decrease from 2.38 eV to 2.08 eV. In contrast, the BE for dissociative adsorption  $\text{Au}_n\text{Y}-2\text{H}$  exhibit a rough decrease from 2.38 to 2.08 eV. More remarkably, for larger clusters ( $n = 4-12$ ), the BE values remain relatively stable, suggesting that these clusters exhibit greater stability after hydrogen release compared to smaller clusters ( $n = 1-3$ ). Furthermore, the higher BE values of  $\text{Au}_n\text{Y}-2\text{H}$  clusters (for  $n = 4-6, 8, 10$  and  $11$ ) compared to their molecularly adsorbed  $\text{Au}_n\text{Y}-\text{H}$  counterparts indicate that  $\text{H}_2$  dissociation following molecular adsorption is more favorable at these sizes. This finding aligns well with the high intrinsic stability of bare  $\text{Au}_6\text{Y}$  clusters.<sup>65</sup> In contrast, for clusters with  $n = 1-3, 7$ , and  $12$ , dissociative adsorption is energetically unfavorable, indicating that  $\text{H}_2$  is more likely to remain in molecular form on these clusters.

## 5. Hydrogen adsorption process

The accumulation of small molecules on a surface is known as adsorption, which can be classified as either physisorption or chemisorption, depending on the nature of the interaction between the adsorbate and the surface. Physisorption is governed by weak van der Waals forces, whereas chemisorption involves the formation of chemical bonds. The adsorption process is influenced by several key factors, including the host cluster, the type of adsorption, and the energy barrier between the adsorbate and the surface. A comprehensive understanding of hydrogen adsorption on  $\text{Au}_n\text{Y}$  ( $n = 2-12$ ) clusters can be obtained through detailed analysis of bonding characteristics, particularly examining the interaction of a hydrogen molecule or two dissociated hydrogen atoms within a specific chemical equilibrium. To gain fundamental insights into the adsorption processes of  $\text{H}_2$  on  $\text{Au}_n\text{Y}$  clusters, we calculate the adsorption energy and H-H bond length for both molecular and dissociative adsorption configurations using the following equations:

$$E_{\text{ads}}(\text{Au}_n\text{Y}-\text{H}_2) = E(\text{Au}_n\text{Y}) + E(\text{H}_2) - E(\text{Au}_n\text{Y}-\text{H}_2) \quad (4)$$

Fig. 2 Comparison of per-atom binding energies (BE, in eV) for  $\text{Au}_n\text{Y}$ ,  $\text{Au}_n\text{Y}-\text{H}_2$ , and  $\text{Au}_n\text{Y}-2\text{H}$  cluster.

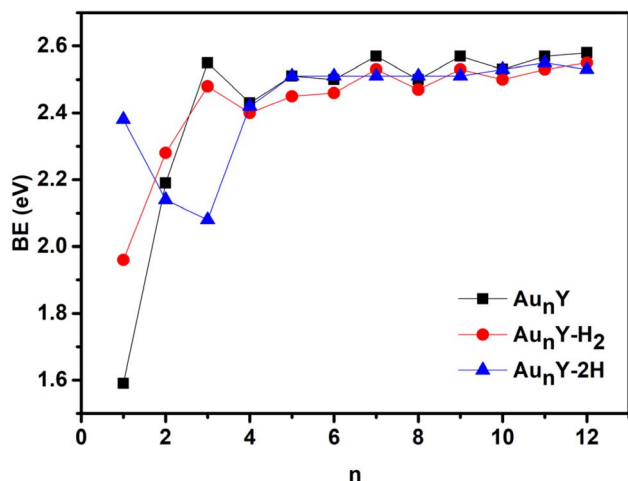


Table 2 Comparison of calculated average binding energies per atom (BE/eV), adsorption energies ( $E_{\text{ads}}$ /eV), and H-H bond length ( $d_{\text{H-H}}$ /Å) for bare  $\text{Au}_n\text{Y}$  clusters, molecular ( $\text{H}_2$ ), and dissociative ( $2\text{H}$ ) hydrogen adsorption on  $\text{Au}_n\text{Y}$  ( $n = 1-12$ )

$n$	BE/eV			$E_{\text{ads}}/\text{eV}$		$d_{\text{H-H}}/\text{\AA}$	
	$\text{Au}_n\text{Y}$	$\text{Au}_n\text{Y}-\text{H}_2$	$\text{Au}_n\text{Y}-2\text{H}$	$\text{Au}_n\text{Y}-\text{H}_2$	$\text{Au}_n\text{Y}-2\text{H}$	$\text{Au}_n\text{Y}-\text{H}_2$	$\text{Au}_n\text{Y}-2\text{H}$
1	1.59	1.96	2.38	0.12	1.81	0.81	3.36
2	2.19	2.28	2.14	0.29	-0.39	0.85	4.16
3	2.55	2.48	2.08	0.15	-2.27	0.75	2.40
4	2.43	2.40	2.42	0.16	0.27	0.76	3.35
5	2.51	2.45	2.51	0.05	0.52	0.83	3.44
6	2.50	2.46	2.51	0.13	0.52	0.76	3.44
7	2.57	2.53	2.51	0.19	0.00	0.75	3.46
8	2.50	2.47	2.51	0.12	0.53	0.74	3.44
9	2.57	2.53	2.51	0.15	-0.11	0.76	3.45
10	2.53	2.50	2.53	0.14	0.56	0.74	3.46
11	2.57	2.53	2.55	0.13	0.40	0.74	3.51
12	2.58	2.55	2.53	0.27	-0.11	0.77	5.54



$$E_{\text{ads}}(\text{Au}_n\text{Y}-2\text{H}) = E(\text{Au}_n\text{Y}) + E(\text{H}_2) - E(\text{Au}_n\text{Y}-2\text{H}) \quad (5)$$

where  $E(\text{Au}_n\text{Y})$ ,  $E(\text{H}_2)$ ,  $E(\text{Au}_n\text{Y}-\text{H}_2)$ , and  $E(\text{Au}_n\text{Y}-2\text{H})$  with  $n = 2-12$  represent the optimized energies of bare  $\text{Au}_n\text{Y}$ , isolated  $\text{H}_2$ , molecularly adsorbed  $\text{Au}_n\text{Y}-\text{H}_2$ , and dissociatively adsorbed  $\text{Au}_n\text{Y}-2\text{H}$ , respectively. The results are summarized in Table 2. A preliminary analysis reveals that, compared to the bond length of isolated  $\text{H}_2$  (0.75 Å),<sup>63</sup> the H-H bond length ( $d_{\text{H-H}}$ ) in  $\text{Au}_n\text{Y}-\text{H}_2$  for  $n = 1-2, 4-6, 9$ , and 12 is significantly elongated, ranging from 0.76 to 0.85 Å. This increase in bond length indicates that  $\text{H}_2$  undergoes activation upon adsorption on these  $\text{Au}_n\text{Y}$  surfaces. In contrast, the distance between hydrogen atoms in  $\text{Au}_n\text{Y}-2\text{H}$  clusters varies from 2.40 Å to 5.54 Å, confirming complete  $\text{H}_2$  dissociation on the cluster surface.

The adsorption energies ( $E_{\text{ads}}$ ) were calculated as the energy difference between the total energy of the reactants ( $\text{Au}_n\text{Y} + \text{H}_2$ ) and the products ( $\text{Au}_n\text{Y}-\text{H}_2$  or  $\text{Au}_n\text{Y}-2\text{H}$ ), as expressed in eqn (4) and (5). In essence,  $E_{\text{ads}}$  represents the minimum energy required to desorb hydrogen from the clusters surface, as listed in Table 2. Comparing the adsorption energy of molecular and dissociative adsorption configurations, we find that all computed  $E_{\text{ads}}$  values for the molecular adsorption ( $\text{Au}_n\text{Y}-\text{H}_2$ ) are positive, ranging from 0.05 to 0.29 eV, indicating that the formation of  $\text{Au}_n\text{Y}-\text{H}_2$  is thermodynamically favorable. For dissociative adsorption ( $\text{Au}_n\text{Y}-2\text{H}$ ),  $E_{\text{ads}}$  values range from -0.39 to 1.81 eV. Notably, the  $E_{\text{ads}}$  values for  $\text{Au}_2\text{Y}-2\text{H}$ ,  $\text{Au}_3\text{Y}-2\text{H}$ ,  $\text{Au}_7\text{Y}-2\text{H}$ ,  $\text{Au}_9\text{Y}-2\text{H}$ , and  $\text{Au}_{12}\text{Y}-2\text{H}$  clusters are negative or zero ( $\text{Au}_7\text{Y}-2\text{H}$ ), suggesting that dissociative adsorption on

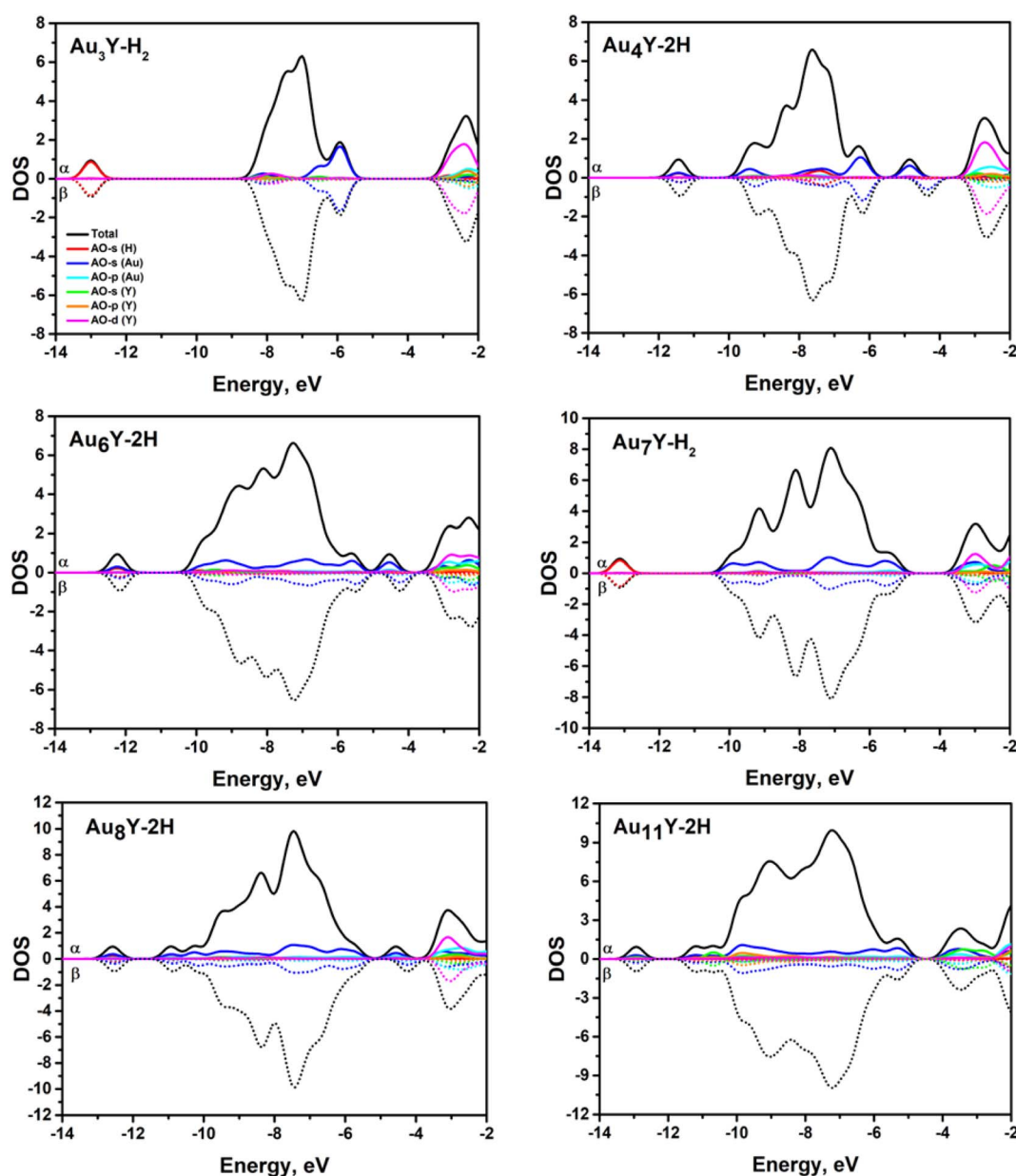


Fig. 3 Density of electronic states (DOS) for the most stable configuration of hydrogen adsorbed on  $\text{Au}_n\text{Y}$  ( $n = 3-4, 6-8$ , and 11) clusters.

these clusters is endothermic and unlikely to occur spontaneously. This phenomenon can be attributed to the decrease in average binding energy for these clusters compared to their bare counterparts ( $\text{Au}_2\text{Y}$ ,  $\text{Au}_3\text{Y}$ ,  $\text{Au}_7\text{Y}$ ,  $\text{Au}_9\text{Y}$ , and  $\text{Au}_{12}\text{Y}$ ). In contrast,  $\text{H}_2$  dissociation on  $\text{Au}_n\text{Y}-2\text{H}$  ( $n = 1, 4-6, 8, 10$  and  $11$ ) clusters is exothermic and thermodynamically favorable, as evidenced by the higher computed  $E_{\text{ads}}$  values for dissociative adsorption  $\text{Au}_n\text{Y}-2\text{H}$  compared to molecular adsorption  $\text{Au}_n\text{Y}-\text{H}_2$ . It is worth mentioning that molecular adsorption is more likely to occur for  $\text{Au}_n\text{Y}$  with  $n = 2, 3, 7, 9$ , and  $12$ . On the other hand, when  $\text{H}_2$  approaches the surface of  $\text{Au}_n\text{Y}$  ( $n = 1, 4-6, 8, 10$ , and  $11$ ), it first undergoes molecular adsorption. Subsequently, the hydrogen molecule may either desorb from the cluster surface or eventually undergo dissociation, forming strong atomic bonds with the cluster. The detachment of adsorbed  $\text{H}_2$  molecules may dominate in certain cases, as it is influenced not only by adsorption energy but also by the activation barrier of the transition state and multiple intermediate states before reaching a stable dissociative configuration. A more detailed discussion on this topic will be provided in the following sections.

To better understand the binding nature of adsorbed hydrogen molecules on  $\text{Au}_n\text{Y}$  ( $n = 1-12$ ) clusters, we analyzed the densities of electronic states (DOS) for both molecular and dissociative adsorption configurations. The DOS results for several thermodynamically preferred configurations are illustrated in Fig. 3, while data for other species are provided in the ESI.†<sup>46</sup> For each DOS graph in Fig. 3, both partial and total DOS are depicted. Solid lines represent spin-up ( $\alpha$ ) states, whereas dashed lines correspond to spin-down ( $\beta$ ) states. As shown in Fig. 3, for species where physisorption is energetically feasible ( $\text{Au}_n\text{Y}-\text{H}_2$  with  $n = 3, 7$ ), most of electronic states associated with adsorbed hydrogen (red line) appear on the low-energy side (below  $-13.0$  eV), indicating a weak interaction between hydrogen and the cluster surface. These states are clearly separated from the electronic states of the cluster, which extend beyond  $-8.5$  eV for  $n = 3$  and  $-10.5$  eV for  $n = 7$ , respectively. This indicates that the chemical bonding contribution between hydrogen and these clusters is relatively weak or potentially negligible. This observation aligns well with previous studies on molecular hydrogen adsorption in transition metal-doped  $\text{Ag}_n$  ( $n = 2-12$ ),  $\text{Au}_{10}^{2+}$  or even  $[\text{Mo}_{13}\text{S}_{13}]^{2-}$  clusters, where van der Waals forces and/or electrostatic interactions between charge-polarized hydrogen molecules and cluster surfaces were identified as the dominant bonding mechanisms.<sup>42,43,67</sup>

The DOS for dissociative configurations of  $\text{Au}_n\text{Y}-2\text{H}$  ( $n = 4, 6, 8, 11$ ) in Fig. 3, along with additional results for  $n = 1, 5$ , and  $10$  provided in the ESI,†<sup>46</sup> exhibit a distinctly different behavior compared to their molecular adsorption counterparts. Remarkably, an analysis of the DOS spectra reveals an intriguing observation: the expected low-energy contribution of  $s\text{-H}_2$  states, typically associated with molecular adsorption, disappears entirely in the dissociative configurations. Instead, the electronic states of hydrogen extend significantly toward higher energies, generally exceeding  $-10.5$  eV, with the exception of  $n = 1$  and  $4$ , where they remain above  $-8.5$  eV. This shift is attributed to the strong hybridization of valence  $s$ -electrons

from hydrogen atoms with those of the cluster, leading to the formation of stable chemical bonds within these systems. Additionally, in most cases, the highest occupied molecular orbital (HOMO) states are predominantly contributed by  $s\text{-Au}$  orbitals, further reinforcing the role of gold in hydrogen dissociation mechanisms. It is important to note that bonding interactions are often accompanied by electron transfer processes. Therefore, these findings support the earlier conclusion that hydrogen adsorption preferentially occurs at Au surface sites.

To gain deeper insight into the dissociative hydrogen adsorption mechanism in  $\text{Au}_n\text{Y}-2\text{H}$  ( $n = 1, 4-6, 8, 10$  and  $11$ ) clusters, we computed possible reaction pathways linking the reactants to their final dissociated states for single  $\text{H}_2$  adsorption and dissociation. Fig. 4 presents the hydrogen dissociation pathways for  $\text{Au}_4\text{Y}$ ,  $\text{Au}_6\text{Y}$ ,  $\text{Au}_8\text{Y}$ , and  $\text{Au}_{11}\text{Y}$  clusters, while additional data for  $\text{AuY}$ ,  $\text{Au}_5\text{Y}$ , and  $\text{Au}_{10}\text{Y}$  are provided in the ESI.†<sup>46</sup> The activation energy for hydrogen dissociation was determined by incorporating zero-point energy corrections and electronic energy values of the reactants and transition states (TS). Specifically, the activation barrier is quantified as the total energy difference between the reactants and the TS, offering insights into the feasibility of hydrogen dissociation on  $\text{Au}_n\text{Y}$  clusters.

As shown in Fig. 4, the interaction of  $\text{H}_2$  with  $\text{Au}_4\text{Y}$  clusters follows a stepwise dissociative adsorption mechanism. Initially, molecular hydrogen preferentially binds to the transition metal atom Y with an adsorption energy of  $-0.16$  eV. However, to activate the adsorbed  $\text{H}_2$ , the  $\text{Au}_4\text{Y}$  clusters must overcome two activation barriers ( $\text{TS}_1$  and  $\text{TS}_2$ ) associated with the dissociation process. The first activation barrier is  $0.78$  eV ( $+0.62$  eV relative to the reactants at  $-0.16$  eV), while the second barrier is  $0.73$  eV ( $+0.60$  eV relative to the intermediate state  $\text{I}_1$  at  $-0.13$  eV). Before reaching the final dissociated product ( $\text{Au}_4\text{Y}-2\text{H}$ ), the system passes through another intermediate state ( $\text{I}_2$ ) at  $0.44$  eV. The final state, where the  $\text{H}_2$  molecule is fully dissociated, has an energy of  $-0.27$  eV, which is  $0.11$  eV lower than the reactant state ( $-0.16$  eV), indicating a thermodynamically favorable process. A similar dissociative adsorption mechanism is observed for  $\text{Au}_n\text{Y}$  clusters with  $n = 1, 5, 6, 8, 10$  and  $11$ ). As detailed in the ESI,† the  $\text{Au}_5\text{Y}$  cluster follows a comparable pathway. The activation barrier for  $\text{H}_2$  dissociation ( $\text{TS}_1$ ) is  $0.75$  eV ( $+0.70$  eV relative to the reactant state), followed by a significant energy decrease through the intermediate state  $\text{I}_1$  ( $-0.04$  eV). Ultimately, the system reaches a fully dissociated state, forming  $\text{Au}_5\text{Y}-2\text{H}$  with an energy of  $-0.52$  eV, which is  $0.47$  eV lower than both the reactants and the  $\text{H}_2$  adsorption state ( $-0.05$  eV).

Similarly, for  $\text{Au}_6\text{Y}$  cluster,  $\text{H}_2$  activation requires overcoming an energy barrier of  $0.68$  eV for the first transition state ( $\text{TS}_1$ ,  $+0.55$  eV relative to the reactant state). The system then proceeds through an intermediate state ( $\text{I}_1$ ) with a slight energy drop of  $0.21$  eV before reaching the final dissociated state, which is  $0.39$  eV lower than the initial molecular adsorption state ( $-0.13$  eV). A similar dissociative adsorption pattern is observed for  $\text{AuY}$ ,  $\text{Au}_8\text{Y}$  and  $\text{Au}_{10}\text{Y}$  clusters. The activation barriers for the first transition state are  $0.48$  eV,  $0.74$  eV and



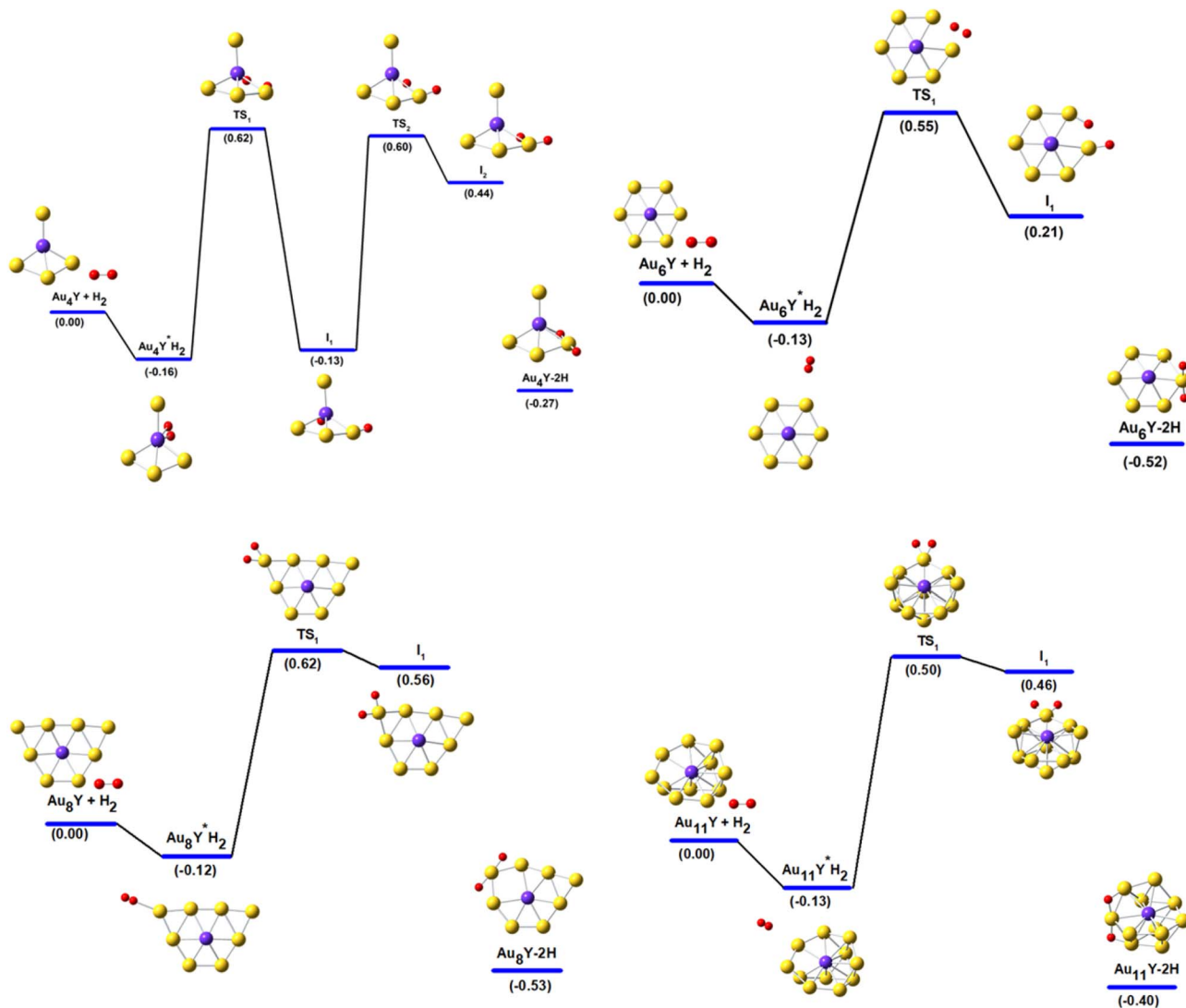


Fig. 4 Computed reaction pathways and relative energies (in eV) for H<sub>2</sub> dissociative adsorption on Au<sub>4</sub>Y, Au<sub>6</sub>Y, Au<sub>8</sub>Y, and Au<sub>11</sub>Y clusters. Intermediates and transition states are denoted as I<sub>i</sub> and TS<sub>i</sub>, respectively. Yellow, purple, and red spheres represent Au, Y, and H atoms, respectively.

0.80 eV for the transition state TS<sub>1</sub> (0.36 eV, 0.62 eV and 0.66 eV relative to the reactant state), respectively. Before reaching their final stable states, Au<sub>8</sub>Y and Au<sub>10</sub>Y pass through intermediate states with energy drops of +0.56 eV and -0.28 eV, respectively. In contrast, the Au<sub>4</sub>Y cluster reaches its final state at -1.81 eV without any intermediate states. Interestingly, the Au<sub>11</sub>Y cluster exhibits the lowest activation energy for H<sub>2</sub> dissociation among the studied clusters, with a barrier of 0.63 eV for TS<sub>1</sub> (+0.50 eV). The system transitions through an intermediate state (I<sub>1</sub>, +0.46 eV) before fully dissociating into the most stable state, Au<sub>11</sub>Y-2H, at -0.40 eV (0.27 eV lower than the reactants and the H<sub>2</sub> adsorption state, -0.13 eV). Comparing the activation barriers for H<sub>2</sub> dissociation across Au<sub>5</sub>Y (0.75 eV), Au<sub>6</sub>Y (0.68 eV), Au<sub>8</sub>Y (0.74 eV), Au<sub>10</sub>Y (0.80 eV), and Au<sub>11</sub>Y (0.63 eV), it is evident that the Au<sub>10</sub>Y cluster exhibits the highest energy barrier at 0.8 eV, suggesting a more challenging dissociative adsorption process compared to other clusters.

Although the final products of H<sub>2</sub> dissociation on the Au<sub>4</sub>Y, Au<sub>5</sub>Y, Au<sub>8</sub>Y, and Au<sub>10</sub>Y clusters exhibit relatively low energies of -0.27 eV, -0.52 eV, -0.53 eV, and -0.56 eV, respectively, the activation energy barriers for H<sub>2</sub> dissociation remain relatively high at 0.78 eV, 0.75 eV, 0.74 eV, and 0.80 eV. This suggests that while H<sub>2</sub> adsorption is energetically favorable for these clusters, dissociative adsorption is unlikely to occur without external stimuli, particularly for the Au<sub>4</sub>Y and Au<sub>10</sub>Y clusters. For the Au<sub>4</sub>Y cluster, the activation barrier for adsorption-dissociation is relatively low at 0.48 eV, and the final product has a significantly low energy of -1.81 eV. However, as previously mentioned, this cluster is known to be unstable, limiting its potential for real-world catalytic applications. In contrast, the Au<sub>6</sub>Y and Au<sub>11</sub>Y clusters exhibit the lowest activation energy barriers of 0.68 eV and 0.63 eV, respectively. This indicates that Au<sub>6</sub>Y and Au<sub>11</sub>Y could serve as promising superatoms for hydrogen storage applications.



## 6. Conclusion

In summary, we have systematically investigated the interaction of H<sub>2</sub> with small-sized Au<sub>n</sub>Y clusters ( $n = 1-12$ ) using DFT calculations. Our findings highlight the crucial role of transition metal dopants in stabilizing cluster geometries and enhancing reactivity toward hydrogen adsorption and dissociation. We identified that the geometric structure, relative electronegativity of the dopant atom, and atomic coordination number ( $N$ ) are key factors governing the preferred adsorption configurations. The Y atom exhibits a strong preference for high-coordination sites, favoring surface positions in smaller clusters ( $n = 1-5$ ) before becoming gradually encapsulated by Au atoms in larger clusters ( $n = 6-12$ ). Due to its relatively lower electronegativity ( $\chi_Y = 1.22$ ) compared to gold ( $\chi_{Au} = 2.54$ ), H<sub>2</sub> adsorption initially favors direct interaction with Y. However, as cluster size increases ( $n = 7-12$ ) and the Y dopant attains higher coordination numbers, hydrogen adsorption shifts toward Au surface sites or Au–Au bridge positions. Notably, our results also indicate that molecular hydrogen adsorption does not significantly alter the structure of bare Au<sub>n</sub>Y clusters (except for Au<sub>5</sub>Y). The adsorption of two dissociated hydrogen atoms preferentially occurs at Au–Au or Au–Y bridge sites, depending on cluster size and atomic arrangement. Average binding energy calculations reveal that the hydrogen adsorption on Au<sub>n</sub>Y ( $n = 4-12$ ) generally maintains higher relative stability compared to their bare counterparts. Importantly, dissociative adsorption is unlikely to occur for Au<sub>n</sub>Y ( $n = 1, 4-5, 8, 10$ ) clusters without external stimulation due to high energy barriers, while for  $n = 2, 3, 7$ , and  $9$  molecular adsorption is energetically preferred. Au<sub>6</sub>Y and Au<sub>11</sub>Y clusters exhibit remarkable stability and barrier-free dissociation pathways, making them promising candidates for hydrogen dissociation and storage applications. Our results provide a deeper insight into the interaction between H<sub>2</sub> and Au<sub>n</sub>Y clusters, including the adsorption site preference, the stability of adsorption configuration, the effect of impurities, possible reaction pathways, and energy barriers involved with this process. This work offers valuable guidance for future experiments and contributes to the development of alloys metal clusters as fundamental components in the design of innovative nanostructured materials for hydrogen catalysis and storage applications.

## Data availability

The data supporting this article have been included as part of the ESI.†

## Author contributions

Nguyen Van Dang: conceptualization, data curation, formal analysis, investigation, validation, writing – original draft, review and editing. Ngo Thi Lan: conceptualization, funding acquisition, data curation, validation, writing – original draft, review and editing. Nguyen Thi Mai: data curation, formal analysis, investigation, validation, review and editing. Son Tung Ngo: data curation, formal analysis, investigation, validation,

review and editing. Phung Thi Thu: data curation, formal analysis, investigation, validation. Nguyen Thanh Tung: conceptualization, data curation, formal analysis, investigation, validation, writing – original draft, review and editing.

## Conflicts of interest

There are no conflicts to declare.

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