

## Nature of Invention: Process design

**Applicant:** BCG

**Inventors:** Naman Goyal, Avi Gupta and Priyanka

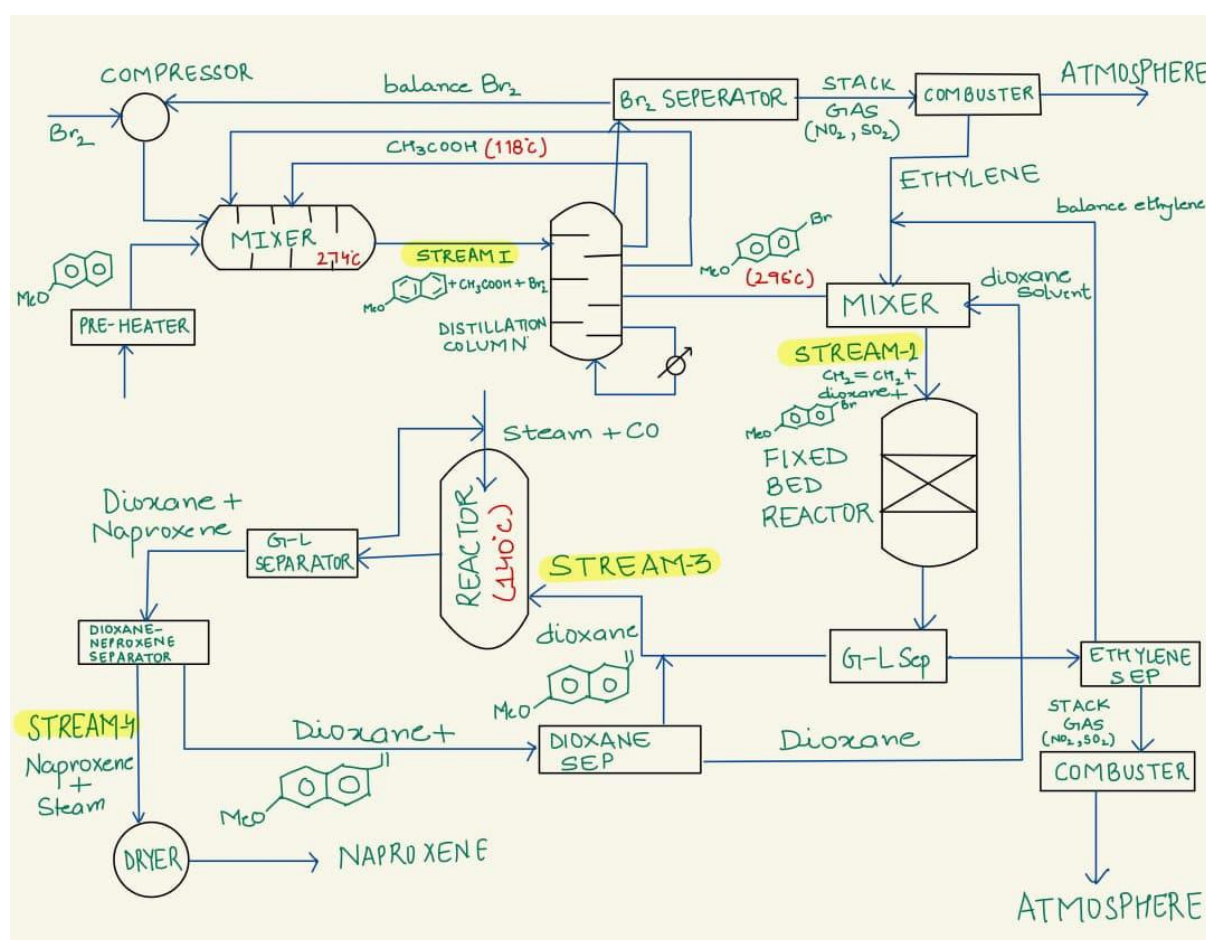
**Chemical Formula:** C<sub>14</sub>H<sub>14</sub>O<sub>3</sub>

**Chemical Name:** (S)-6-Methoxy- $\alpha$ -methyl-2-naphthaleneacetic acid

**Process Title:** Production of Naproxen from Methoxy Naphthylene

### Process Description:

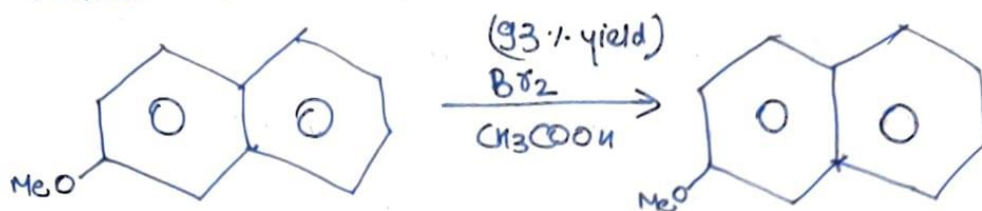
a. Block Diagram:



Catalysts for Fixed Bed Reactor are:  $\text{Pd}(\text{OAc})_2$  and  $\text{PPh}_3$

b. Material Balance:

Material Balance:



Assumption: We have  $x$  Mole of Methoxynaphthalene  
 ( $\text{C}_{10}\text{H}_7\text{C}_1\text{H}_9\text{O}$ )

After Recycled Mole:  $y$  moles

After Recycling we have  $(x+y)$  moles of  $\text{C}_{11}\text{H}_{10}\text{O}$

$$\Rightarrow 0.07(x+y) = 0.93y$$

$$\frac{0.07x}{0.93} = y$$

$$y = 0.075x$$

Total moles in ~~Star~~ Reactor/Mixer =  $1.075x$

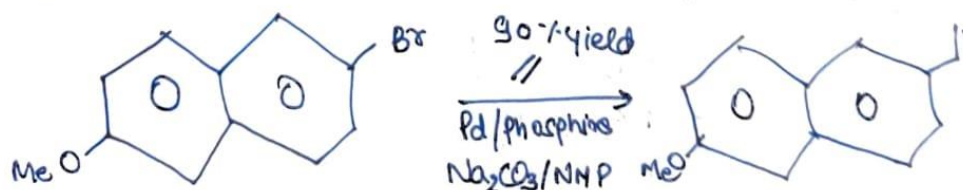
We need 1 mol of  $\text{Br}_2$  for react with 1 mol of  $\text{C}_{11}\text{H}_{10}\text{O}$  so we need total  $1.075x$  mole of  $\text{Br}_2$  in Reactor

$$\begin{aligned}\text{Recycled/Purge of } \text{Br}_2 \text{ mole} &= 1.075x \times 0.07 \\ &= 0.075x \text{ moles}\end{aligned}$$

Initial moles of  $\text{Br}_2$  Required =  $x$  moles

$$\begin{aligned}\text{Total Moles of 2-Bromo 6-Methoxy Naphthalene} &= (x+y) \times 0.93 \\ &= 1.075x \times 0.93 \\ &= x \text{ moles}\end{aligned}$$

Heck  $\text{R}^{\text{H}}$ :



Total Moles after recycling of 2-Bromo-6-Methoxy Naphthalene going into reactor =  $(x+y)$  moles

where  $y$  is Total moles recycled

$$0.1(x+y) = y$$

$$\frac{x}{9} = y$$

Total Moles of 2-Bromo-6-Methoxy Naphthalene in Reactor =  $\frac{10x}{9}$  moles

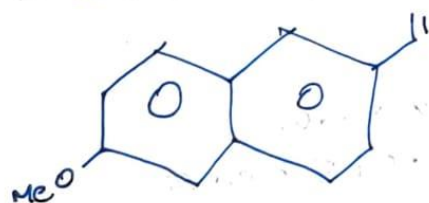
Required Moles of Ethylene in Reactor =  $\frac{10x}{9}$  moles

Unreacted/Recycled Moles of Ethylene =  $\frac{10x}{9} \times \frac{1}{10} = \frac{x}{9}$  moles

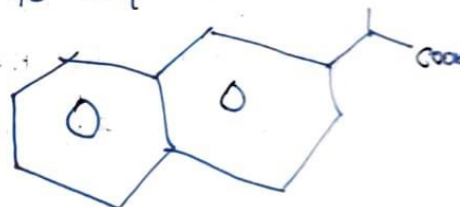
Required Moles of Ethylene =  $x$  moles

Now final part is to convert

Eth-1-(6-Methoxy Naphthalene) to Naproxene



$\xrightarrow[\text{(86.7\% yield)}]{\text{Co/H}_2\text{O}}$



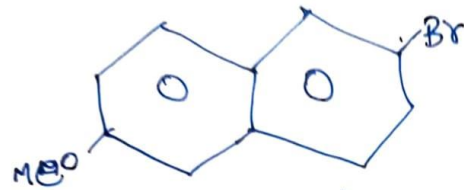
Total Moles of Eth-1-(6-Methoxy Naphthalene) in Reactor =  $(x+y)$  moles where  $y$  is recycled moles

We need 1 mol of CO for the rxn with 1 mol of Eth-1-(6-Methoxy Naphthalene)

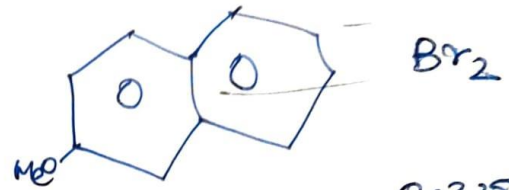
$$\Rightarrow 0.14(x+y) = y$$

$$\frac{0.14x}{0.86} = y$$

Stream I:  $\text{CH}_3\text{COOH}$  is acting as solvent



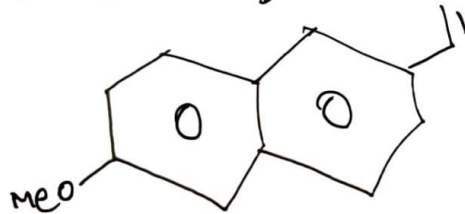
$$\begin{aligned} & 4.34 \text{ kmol/day} \\ &= 0.18 \text{ kmol/hr} \\ &= 42.85 \text{ kg/hr} \end{aligned}$$



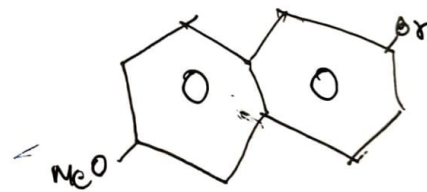
$$\begin{aligned} & 0.3255 \text{ kmol/day} \\ &= 0.013 \text{ kmol/hr} \\ &= 2.054 \text{ kg/hr} \end{aligned}$$

$$\begin{aligned} & 0.3255 \text{ kmol/day} \\ &= 0.013 \text{ kmol/hr} \\ &= 2.07 \text{ kg/hr} \end{aligned}$$

Stream 2: Dioxane is acting solvent



$$\begin{aligned} & 4.34 \text{ kmol} \\ &= 0.18 \text{ kmol/hr} \\ &= 33.12 \text{ kg/hr} \end{aligned}$$



$$\begin{aligned} & 0.482 \text{ kmol/day} \\ &= 0.02 \text{ kmol/hr} \\ &= 4.78 \text{ kg/hr} \end{aligned}$$

$$\begin{aligned} & \text{Ethylene} \\ & 0.482 \text{ kmol} \end{aligned}$$

$$\begin{aligned} &= 0.02 \text{ kmol/hr} \\ &= 0.562 \text{ kg/hr} \end{aligned}$$

c. Capital Cost

#### References:


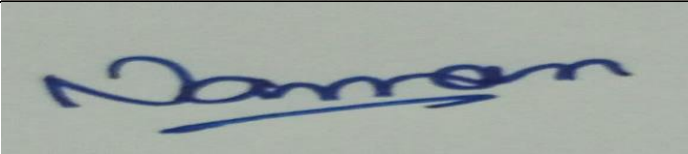

1. <http://www.matche.com/equipcost/Reactor.html>
2. <https://www.sciencedirect.com/topics/engineering/fixed-bed-reactor>
3. <https://www.sciencedirect.com/topics/engineering/distillation-column>
4. <https://www.mdpi.com/2073-4344/7/9/267>

#### List the contributions of each author:

- Avi Gupta and Naman Goyal designed the Block Diagram
- Naman Goyal did the Material Balance

- Avi Gupta and Priyanka calculated Capital Cost

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