

CH365 Chemical Engineering Thermodynamics

Lesson 36

Fugacity & Fugacity Coefficients: Species in Solution

Block 6 – Solution Thermodynamics

Lesson 35 Recap - Fugacity

$$G_i^{\text{ig}} = \Gamma_i(T) + RT \ln P \quad (\text{Eq. 10.28})$$

$$G_i = \Gamma_i(T) + RT \ln f_i \quad (\text{Eq. 10.31})$$

$$G_i - G_i^{\text{ig}} = RT \ln \frac{f_i}{P}$$

$$G_i^R = RT \ln \phi_i \quad (\text{Eq. 10.33})$$

$$\phi_i \equiv \frac{f_i}{P} \quad (\text{Eq. 10.34})$$

- units of pressure; “escaping tendency”

- *residual Gibbs energy*:

$$G_i^R = G_i - G_i^{\text{ig}} \quad (\text{Eq. 6.41})$$

- for ideal gases:

$$f_i^{\text{ig}} = P \quad (\text{Eq. 10.32})$$

- The residual Gibbs energy of component *i* in a mixture is a function of fugacity

$$\frac{G^R}{RT} = \int_0^P (Z - 1) dP \quad (\text{Eq. 6.49})$$



$$\frac{G_i^R}{RT} = \ln \phi_i \quad (\text{Eq. 10.33})$$



$$\ln \phi_i = \int_0^P (Z_i - 1) dP \quad (\text{Eq. 10.35})$$

Cubic Equations of State:
Derived in Chapter 6

$$\beta_i = \Omega \frac{P_{r_i}}{T_r} \quad (\text{Eq. 3.50})$$

$$q_i = \frac{\Psi \alpha}{\Omega T_{r_i}} \quad (\text{Eq. 3.51})$$

$$I_i = \frac{1}{\sigma - \epsilon} \ln \left(\frac{Z_i + \sigma \beta}{Z_i + \epsilon \beta} \right) \quad (\text{Eq. 13.72})$$

$$\ln \phi_i = Z_i - 1 - \ln(Z_i - \beta_i) - q_i I_i \quad (\text{Eq. 13.85})$$

Vapor/Liquid Equilibrium

$$(Eq. 10.31) \rightarrow G_i^{\text{vapor}} = \Gamma_i(T) + RT \ln f_i^{\text{vapor}} \quad (Eq. 10.37)$$

$$\leftarrow G_i^{\text{liquid}} = \Gamma_i(T) + RT \ln f_i^{\text{liquid}} \quad (Eq. 10.38)$$

$$G_i^{\text{vapor}} - G_i^{\text{liquid}} = RT \ln \frac{f_i^{\text{vapor}}}{f_i^{\text{liquid}}}$$

(Equilibrium, Eq. 6.83)

$$G_\alpha = G_\beta \quad \rightarrow \quad G_i^{\text{vapor}} = G_i^{\text{liquid}} \quad \ln(1) = 0$$

$$f_i^{\text{vapor}} = f_i^{\text{liquid}} = f_i^{\text{sat}} \quad (Eq. 10.39)$$

When the escaping tendency is the same for the two phases, they are in equilibrium.
When the escaping tendency of a species is higher in one phase than another, that species will tend to transfer to the phase where fugacity is lower.

$$\phi_i \equiv \frac{f_i}{P} \quad (Eq. 10.34)$$

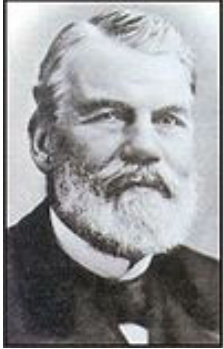
$$\phi_i^{\text{sat}} \equiv \frac{f_i^{\text{sat}}}{P_i^{\text{sat}}} \quad (Eq. 10.40)$$

$$\phi_i^{\text{vapor}} = \phi_i^{\text{liquid}} = \phi_i^{\text{sat}} \quad (Eq. 10.41)$$

For pure species, coexisting liquid and vapor phases are in equilibrium when they have the same temperature, pressure, fugacity, and fugacity coefficient

Raoult's Law

Wikipedia



Francois-Marie Raoult,
1830-1901

"General Law of the Vapor Pressure of Solvents," in the French Journal, *Comptes Rendus* (May 23, 1887)

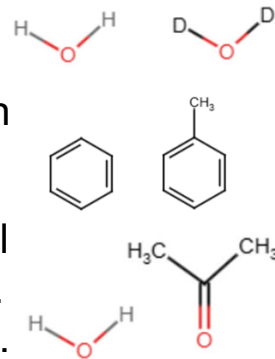
The partial pressure of each component of an ideal mixture of liquids is equal to the vapor pressure of the pure component multiplied by its mole fraction in the liquid mixture.

$$y_i \cdot P = x_i \cdot P_i^{\text{sat}}$$

Partial pressure of each component

Vapor pressure multiplied by liquid solution mole fraction

Raoult, 1887



Phenomenological; assumes ideal behavior; ideal behavior in a solution is when IMFs between unlike molecules are equal to IMFs between similar molecules.

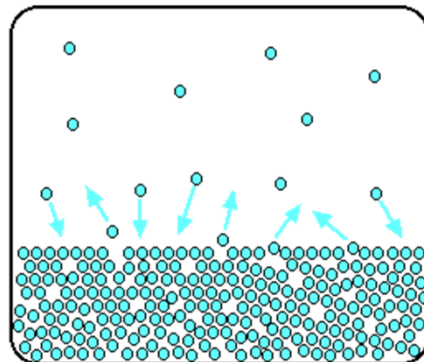
The more similar the components are, the more their behavior approaches ideal solution (Raoult's Law) behavior. For example, H_2O and D_2O follow RL exactly. Benzene and toluene follow it very closely. Acetone and water do not follow RL.

Vapor pressure depression – one of the "colligative properties"

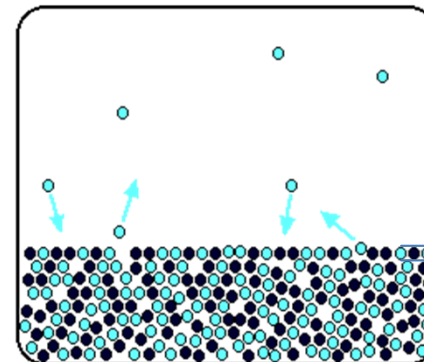
IMFs:

If $\bullet\text{---}\bullet = \text{---}\bullet\text{---}\bullet$, then $x_i P_i^{\text{sat}}$ accurately describes reduction

If $\bullet\text{---}\bullet \neq \text{---}\bullet\text{---}\bullet$, then we introduce the "activity coefficient" γ_i .



Equilibrium between volatile molecules in liquid and gas phases



Equilibrium between volatile molecules in liquid and gas phases with solute present

The fraction of sites in the surface layer occupied by " \bullet " is equal to the mole fraction of " \bullet " in the bulk solution. The " \bullet " reduce the " $\text{---}\bullet$ " in the surface layer.

$A(\text{liq}) \rightleftharpoons A(\text{vap})$

Raoult's Law K-values: $\frac{y_i}{x_i} = \frac{P_i^{\text{sat}}}{P} = K_i$

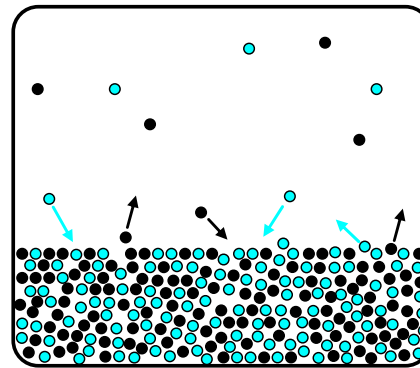
with unequal IMFs:

$x_i P_i^{\text{sat}}$ is replaced with $\gamma_i x_i P_i^{\text{sat}}$

$$\Rightarrow \frac{y_i}{x_i} = \frac{\gamma_i P_i^{\text{sat}}}{P} = K_i$$

Raoult's Law

When both molecules can escape to the vapor, the behavior of each substance can be described by Raoult's Law, and the vapor pressure of each substance is reduced by the mole fraction of the substance in the liquid.



● volatile

● volatile

$$y_{\bullet} \cdot P = x_{\bullet} P_{\bullet}^{\text{sat}}$$

$$y_{\bullet} \cdot P = x_{\bullet} P_{\bullet}^{\text{sat}}$$

$$y_{\bullet} \cdot P + y_{\bullet} \cdot P = P_{\text{TOT}} = x_{\bullet} P_{\bullet}^{\text{sat}} + x_{\bullet} P_{\bullet}^{\text{sat}}$$

Vapor pressure given by P^{sat}
from the Antoine Equation

$$P_{\bullet}^{\text{sat}} = e^{A_{\bullet} - \frac{B_{\bullet}}{T+C_{\bullet}}}$$

$$P_{\bullet}^{\text{sat}} = e^{A_{\bullet} - \frac{B_{\bullet}}{T+C_{\bullet}}}$$

Equilibrium Condition

$$\mu_i^\alpha = \mu_i^\beta = \dots = \mu_i^\pi \quad (\text{Eq. 10.6, page 361})$$

$$f_i^{\text{vapor}} = f_i^{\text{liquid}} = f_i^{\text{sat}} \quad (\text{Eq. 10.39, page 377})$$

$$\phi_i^{\text{vapor}} = \phi_i^{\text{liquid}} = \phi_i^{\text{sat}} \quad (\text{Eq. 10.41, page 378})$$

For pure species, coexisting liquid and vapor phases are in equilibrium when they have the same temperature, pressure, Gibbs energy, chemical potential, fugacity and fugacity coefficient.

Replace pressures in Raoult's Law with fugacities

$$K_i = \frac{y_i}{x_i} = \frac{P_i^{\text{sat}}}{P} \quad \Rightarrow \quad K_i = \frac{y_i}{x_i} = \frac{f_i^{\text{sat}}}{f_i}$$

Raoult's Law K-values

$$\begin{aligned} y_i P &= x_i P_i^{\text{sat}} \\ y_i f_i &= x_i f_i^{\text{sat}} \\ \frac{P_i}{P} f_i &= x_i f_i^{\text{sat}} \\ \frac{f_i}{P} P_i &= x_i f_i^{\text{sat}} \\ \phi_i P_i &= x_i f_i^{\text{sat}} \\ \phi_i y_i P &= x_i f_i^{\text{sat}} \end{aligned}$$

$$K_i = \frac{y_i}{x_i} = \frac{\gamma_i P_i^{\text{sat}}}{P} \quad \Rightarrow \quad K_i = \frac{y_i}{x_i} = \frac{\gamma_i f_i^{\text{sat}}}{f_i}$$

Modified Raoult's Law K-values

For solutions, multiple phases at the same T and P are in equilibrium when the fugacity of each species is the same.

fugacity of species i in
solution in phase α

$$\hat{f}_i^\alpha$$

$$\text{in solution, } \hat{f}_i^\alpha = \hat{f}_i^\beta = \dots = \hat{f}_i^\pi \quad (\text{Eq. 10.47, page 382})$$

$$\bar{G}_i^R = RT \ln \hat{\phi}_i \quad (\text{Eq. 10.51})$$

$$\hat{\phi}_i = \frac{\hat{f}_i}{y_i P} \quad (\text{Eq. 10.52, page 383})$$

Fugacity of Liquids

by Equation 10.34

Poynting equation:

Derived on pp. 378-379

$$f_i = \underbrace{\phi_i^{\text{sat}}}_{f_i^{\text{sat}}} P_i^{\text{sat}} \exp \left[\frac{V_i^l (P - P_i^{\text{sat}})}{RT} \right] \quad (\text{Eq. 10.44, page 379})$$

$$K_i = \frac{y_i}{x_i} = \frac{f_i^{\text{sat}}}{f_i}$$

$$\therefore \frac{f_i}{f_i^{\text{sat}}} = \exp \left[\frac{V_i^l (P - P_i^{\text{sat}})}{RT} \right] = \text{Poynting factor}$$

P_i^{sat} saturation vapor pressure

V_i^l liquid molar volume, assumed constant

$$\phi_i^{\text{sat}} = \frac{f_i^v(P_i^{\text{sat}})}{P_i^{\text{sat}}} \quad (\text{Eq. (A), page 378})$$

(vapor-phase fugacity coefficient)

read as “fugacity evaluated at saturation pressure,” not “multiplied by”

$$\ln \phi_i = \int_0^P (Z_i - 1) dP \quad \Rightarrow \quad \ln \phi_i^{\text{sat}} = \int_0^{P_i^{\text{sat}}} (Z_i^v - 1) dP$$

(Eq. 10.35)

(Eq. 10.35 and 10.42)

Questions?