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<div>5</div> <div>Boron</div> <div>10.81</div> <div>Ceramics</div> <div>B</div> <div>Boron is the 5th element but essential for plants and harder than diamond, used in ceramics, glass, and as neutron absorber</div>	<div>6</div> <div>Carbon</div> <div>12.011</div> <div>Steel</div> <div>C</div> <div>Carbon is highest for 2 hundred million compounds, used in steel production and as graphite</div>	<div>7</div> <div>Nitrogen</div> <div>14.007</div> <div>Fertilizer</div> <div>N</div> <div>Nitrogen makes up 78% of Earth's atmosphere but is used in fertilizers and as liquid nitrogen coolant</div>	<div>8</div> <div>Oxygen</div> <div>15.999</div> <div>Breathing</div> <div>O</div> <div>Oxygen is Earth's most abundant element at 46% of atmosphere, essential for breathing and used in steel production</div>
<div>9</div> <div>Fluorine</div> <div>18.998</div> <div>Toothpaste</div> <div>F</div> <div>Fluorine is the most reactive element that can corrode glass, used in toothpaste and water fluoridation</div>	<div>10</div> <div>Neon</div> <div>20.18</div> <div>Lighting</div> <div>Ne</div> <div>Neon produces the most intense light discharge creating neon signs and as inert gas in lighting</div>	<div>11</div> <div>Sodium</div> <div>22.99</div> <div>Salt</div> <div>Na</div> <div>Sodium lamps are so efficient that one can outshine 100 incandescent bulbs, used in table salt and street lighting</div>	<div>12</div> <div>Magnesium</div> <div>24.305</div> <div>Flares</div> <div>Mg</div> <div>Magnesium burns with 3000K white light so bright it can be used in flares, alloys, and as dietary supplement</div>
<div>13</div> <div>Aluminum</div> <div>26.982</div> <div>Cans</div> <div>Al</div> <div>Aluminum was worth more than gold until 1890s before being used in beverage cans, foil, and aircraft parts</div>	<div>14</div> <div>Silicon</div> <div>28.086</div> <div>Chips</div> <div>Si</div> <div>Silicon makes up 27% of Earth's crust and enabled the computer age, used in computer chips, glass, and construction</div>	<div>15</div> <div>Phosphorus</div> <div>30.974</div> <div>Matches</div> <div>P</div> <div>White phosphorus glows green in darkness but is essential for life, used in fertilizers, matches, and DNA structure</div>	<div>16</div> <div>Sulfur</div> <div>32.06</div> <div>Gunpowder</div> <div>S</div> <div>Sulfur is second for forming 30+ allotropes and creates rubber, used in rubber vulcanization and gunpowder</div>
<div>17</div> <div>Chlorine</div> <div>35.45</div> <div>Disinfectant</div> <div>Cl</div> <div>Chlorine has the highest electron affinity (349 kJ/mol), used in pool disinfection and PVC production</div>	<div>18</div> <div>Argon</div> <div>39.948</div> <div>Inert</div> <div>Ar</div> <div>Argon was Earth's first isolated noble gas and makes up 1% of atmosphere, used in welding and incandescent light bulbs</div>	<div>19</div> <div>Potassium</div> <div>39.098</div> <div>Soap</div> <div>K</div> <div>Potassium is so violently reactive it ignites in water, used in fertilizers and soap production</div>	<div>20</div> <div>Calcium</div> <div>40.078</div> <div>Bones</div> <div>Ca</div> <div>Calcium phosphate comprises 70% of bone mass giving strength, used in bones, teeth, and concrete production</div>

<div>21</div> <div>Aerospace</div> <div>Scandium</div> <div>Sc</div> <div>44.956</div> <div>Scandium is paradoxically rarer than many 'rare earth'</div>	<div>22</div> <div>Implants</div> <div>Titanium</div> <div>Ti</div> <div>47.867</div> <div>Titanium has the highest strength-to-weight ratio of all</div>	<div>23</div> <div>Pigments</div> <div>Vanadium</div> <div>V</div> <div>50.942</div> <div>Vanadium can exist in 5 different oxidation states</div>	<div>24</div> <div>Plating</div> <div>Chromium</div> <div>Cr</div> <div>51.996</div> <div>Chromium gives rubies red and emeralds green color</div>
<div>25</div> <div>Electrodes</div> <div>Manganese</div> <div>Mn</div> <div>54.938</div> <div>Manganese nodules carpet vast ocean floors containing</div>	<div>26</div> <div>Construction</div> <div>Iron</div> <div>Fe</div> <div>55.845</div> <div>Iron comprises 32.1% of Earth's total mass with most</div>	<div>27</div> <div>Glass</div> <div>Cobalt</div> <div>Co</div> <div>58.933</div> <div>Cobalt blue glass has been prized for 4000 years and</div>	<div>28</div> <div>Currency</div> <div>Nickel</div> <div>Ni</div> <div>58.693</div> <div>Nickel-iron meteorites delivered much of Earth's</div>
<div>29</div> <div>Wiring</div> <div>Copper</div> <div>Cu</div> <div>63.546</div> <div>Copper naturally kills bacteria and viruses within hours</div>	<div>30</div> <div>Galvanizing</div> <div>Zinc</div> <div>Zn</div> <div>65.38</div> <div>Zinc deficiency causes loss of taste/smell and affects 2 billion</div>	<div>31</div> <div>LEDs</div> <div>Gallium</div> <div>Ga</div> <div>69.723</div> <div>Gallium melts at 29.8°C in hand temperature but boils at</div>	<div>32</div> <div>Transistors</div> <div>Germanium</div> <div>Ge</div> <div>72.63</div> <div>Germanium was predicted by Mendeleev 15 years before</div>
<div>33</div> <div>Preservatives</div> <div>Arsenic</div> <div>As</div> <div>74.922</div> <div>Arsenic has been humanity's poison of choice for over 2000</div>	<div>34</div> <div>Photoconductors</div> <div>Selenium</div> <div>Se</div> <div>78.974</div> <div>Selenium deficiency causes fatal white muscle disease and</div>	<div>35</div> <div>Antiseptic</div> <div>Bromine</div> <div>Br</div> <div>79.904</div> <div>Bromine is the only liquid non-metal but it evaporates</div>	<div>36</div> <div>Windows</div> <div>Krypton</div> <div>Kr</div> <div>83.798</div> <div>Krypton was used in ultra-bright airport runway</div>
<div>37</div> <div>Clocks</div> <div>Rubidium</div> <div>Rb</div> <div>85.468</div> <div>Rubidium ignites spontaneously in air and was</div>	<div>38</div> <div>Fireworks</div> <div>Strontium</div> <div>Sr</div> <div>87.62</div> <div>Strontium-90 fallout creates the brilliant red in fireworks but</div>	<div>39</div> <div>Lasers</div> <div>Yttrium</div> <div>Y</div> <div>88.906</div> <div>Yttrium with barium carbon oxide named YBCO makes the</div>	<div>40</div> <div>Ceramics</div> <div>Zirconium</div> <div>Zr</div> <div>91.224</div> <div>Zirconium is virtually immune to corrosion up to 1270K and</div>

<div>41</div> <div>Engines</div> <div>Nb</div> <div>200.6</div> <div>Niobium</div> <div> </div> <div> <p>Niobium is superconducting below 9K and was originally</p> </div> <div>Used in jet engines and MRI scanners</div>	<div>42</div> <div>Lubricants</div> <div>Mo</div> <div>95.94</div> <div>Molybdenum</div> <div> </div> <div> <p>Molybdenum has the 6th highest melting point at 2896K</p> </div> <div>Used in steel alloys and high-temp lubricants</div>	<div>43</div> <div>Isotope</div> <div>Tc</div> <div>98</div> <div>Technetium</div> <div> </div> <div> <p>Technetium was the first artificially created element</p> </div> <div>Used in medical imaging and as tracer</div>	<div>44</div> <div>Contacts</div> <div>Ru</div> <div>101.07</div> <div>Ruthenium</div> <div> </div> <div> <p>Ruthenium is the scarcest platinum group metal and</p> </div> <div>Used in electrical contacts and hard disks</div>
<div>45</div> <div>Jewelry</div> <div>Rh</div> <div>102.91</div> <div>Rhodium</div> <div> </div> <div> <p>Rhodium is the most expensive precious metal at</p> </div> <div>Used in catalytic converters and jewelry</div>	<div>46</div> <div>Dentistry</div> <div>Pd</div> <div>106.42</div> <div>Palladium</div> <div> </div> <div> <p>Palladium can absorb 900 times its volume in hydrogen</p> </div> <div>Used in catalytic converters and dentistry</div>	<div>47</div> <div>Photography</div> <div>Ag</div> <div>107.87</div> <div>Silver</div> <div> </div> <div> <p>Silver has the highest electrical conductivity of all elements at</p> </div> <div>Used in jewelry, mirrors, and photography</div>	<div>48</div> <div>Panels</div> <div>Cd</div> <div>112.41</div> <div>Cadmium</div> <div> </div> <div> <p>Cadmium red paint was banned after causing severe</p> </div> <div>Used in batteries, pigments, and solar panels</div>
<div>49</div> <div>Displays</div> <div>In</div> <div>114.82</div> <div>Indium</div> <div> </div> <div> <p>Indium is softer than lead and can be scratched with a</p> </div> <div>Used in semiconductors and LCD screens</div>	<div>50</div> <div>Solder</div> <div>Sn</div> <div>118.71</div> <div>Tin</div> <div> </div> <div> <p>Tin produces a distinctive 'tin cry' scream when bent due to</p> </div> <div>Used in solder, cans, and bronze alloys</div>	<div>51</div> <div>Alloys</div> <div>Sb</div> <div>121.76</div> <div>Antimony</div> <div> </div> <div> <p>Fluoroantimonic acid is 10 quintillion times stronger than</p> </div> <div>Used in flame retardants and semiconductors</div>	<div>52</div> <div>Solar</div> <div>Te</div> <div>127.6</div> <div>Tellurium</div> <div> </div> <div> <p>Tellurium-128 has the longest known half-life at 2.2 septillion</p> </div> <div>Used in solar panels and rubber vulcanization</div>
<div>53</div> <div>Medicine</div> <div>I</div> <div>126.91</div> <div>Iodine</div> <div> </div> <div> <p>Iodine deficiency affects 2 billion people causing goiter</p> </div> <div>Used as antiseptic and in photography</div>	<div>54</div> <div>Anesthesia</div> <div>Xe</div> <div>131.29</div> <div>Xenon</div> <div> </div> <div> <p>Xenon is the rarest gas with 90 grams per million kilograms of</p> </div> <div>Used in ion drives and medical anesthesia</div>	<div>55</div> <div>Timekeeping</div> <div>Cs</div> <div>132.91</div> <div>Cesium</div> <div> </div> <div> <p>Caesium is the softest metal and its hydroxide is the</p> </div> <div>Used in atomic clocks and oil drilling</div>	<div>56</div> <div>Imaging</div> <div>Ba</div> <div>137.33</div> <div>Barium</div> <div> </div> <div> <p>Barium compounds create brilliant green fireworks but are</p> </div> <div>Used in X-ray imaging and drilling fluids</div>
<div>57</div> <div>Lighters</div> <div>La</div> <div>138.91</div> <div>Lanthanum</div> <div> </div> <div> <p>Lanthanum remained undiscovered in 'pure' cerium</p> </div> <div>Used in lighter flints and camera lenses</div>	<div>58</div> <div>Polishing</div> <div>Ce</div> <div>140.12</div> <div>Cerium</div> <div> </div> <div> <p>Cerium is the most abundant rare earth comprising 0.006%</p> </div> <div>Used in catalysts and glass polishing</div>	<div>59</div> <div>Engines</div> <div>Pr</div> <div>140.91</div> <div>Praseodymium</div> <div> </div> <div> <p>Praseodymium means 'green twin' creating emerald-green</p> </div> <div>Used in aircraft engines and magnets</div>	<div>60</div> <div>Speakers</div> <div>Nd</div> <div>144.24</div> <div>Neodymium</div> <div> </div> <div> <p>Neodymium creates the strongest permanent magnets</p> </div> <div>Used in powerful permanent magnets</div>

<div>61</div> <div>P</div> <div>Promethium</div> <div>Batteries</div> <div>Pm</div> <div>145</div> <div>Promethium is the only radioactive rare earth and</div> <div>Used in nuclear batteries and research</div>	<div>62</div> <div>Sm</div> <div>Samarium</div> <div>Reactors</div> <div>Sm</div> <div>153.9</div> <div>Samarium magnets work at 350°C and have the highest</div> <div>Used in magnets and cancer treatment</div>	<div>63</div> <div>Eu</div> <div>Europium</div> <div>Phosphors</div> <div>Eu</div> <div>151.96</div> <div>Europium is the softest rare earth and the most reactive,</div> <div>Used in red phosphors for TV screens</div>	<div>64</div> <div>Gd</div> <div>Gadolinium</div> <div>Contrast</div> <div>Gd</div> <div>157.25</div> <div>Gadolinium has the highest magnetic moment and is used</div> <div>Used as MRI contrast agents and neutron capture</div>
<div>65</div> <div>Tb</div> <div>Terbium</div> <div>Green</div> <div>Tb</div> <div>158.93</div> <div>Terbium glows intense green under UV and is essential for</div> <div>Used in green phosphors and magnets</div>	<div>66</div> <div>Dy</div> <div>Dysprosium</div> <div>Lasers</div> <div>Dy</div> <div>162.5</div> <div>Dysprosium becomes strongly magnetic only below -180°C</div> <div>Used in lasers and hard disk drives</div>	<div>67</div> <div>Ho</div> <div>Holmium</div> <div>Devices</div> <div>Ho</div> <div>164.93</div> <div>Holmium possesses the strongest magnetic field of any</div> <div>Used in magnets and medical devices</div>	<div>68</div> <div>Er</div> <div>Erbium</div> <div>Fiber</div> <div>Er</div> <div>167.26</div> <div>Erbium amplifies light in fiber optic cables enabling global</div> <div>Used in fiber optic amplifiers and lasers</div>
<div>69</div> <div>Tm</div> <div>Thulium</div> <div>Sources</div> <div>Tm</div> <div>168.93</div> <div>Thulium is the least abundant rare earth metal and possibly</div> <div>Used in X-ray sources and portable equipment</div>	<div>70</div> <div>Yb</div> <div>Ytterbium</div> <div>Sensors</div> <div>Yb</div> <div>173.05</div> <div>Ytterbium expands 26% during phase transition and is used in</div> <div>Used in lasers and stress gauges</div>	<div>71</div> <div>Lu</div> <div>Lutetium</div> <div>Catalysis</div> <div>Lu</div> <div>174.97</div> <div>Lutetium is the hardest, densest rare earth and was the</div> <div>Used in catalysts and medical imaging</div>	<div>72</div> <div>Hf</div> <div>Hafnium</div> <div>Cutting</div> <div>Hf</div> <div>178.49</div> <div>Hafnium has nearly identical properties to zirconium due to</div> <div>Used in tungsten carbide and nuclear reactor</div>
<div>73</div> <div>Ta</div> <div>Tantalum</div> <div>Capacitors</div> <div>Ta</div> <div>180.95</div> <div>Tantalum is virtually immune to all acids except hydrofluoric at</div> <div>Used in electronics and surgical instruments</div>	<div>74</div> <div>W</div> <div>Tungsten</div> <div>Filaments</div> <div>W</div> <div>183.84</div> <div>Tungsten has the highest melting point at 3695K and</div> <div>Used in light bulb filaments and X-ray tubes</div>	<div>75</div> <div>Re</div> <div>Rhenium</div> <div>Coatings</div> <div>Re</div> <div>186.21</div> <div>Rhenium has the highest boiling point at 5869K and is</div> <div>Used in catalysts and jet engine parts</div>	<div>76</div> <div>Os</div> <div>Osmium</div> <div>Contacts</div> <div>Os</div> <div>190.23</div> <div>Osmium is the densest element at 22.6 g/cm³ and</div> <div>Used in fountain pen tips and electrical contacts</div>
<div>77</div> <div>Ir</div> <div>Iridium</div> <div>Plugs</div> <div>Ir</div> <div>192.22</div> <div>Iridium is the most corrosion-resistant element</div> <div>Used in spark plugs and cancer treatment</div>	<div>78</div> <div>Pt</div> <div>Platinum</div> <div>Refining</div> <div>Pt</div> <div>195.08</div> <div>Platinum is 30 times rarer than gold and catalyzes 20% of all</div> <div>Used in jewelry, catalysts, and electronics</div>	<div>79</div> <div>Au</div> <div>Gold</div> <div>Bullion</div> <div>Au</div> <div>196.97</div> <div>Gold is so chemically inert it never tarnishes and has been</div> <div>Used in jewelry, electronics, and dentistry</div>	<div>80</div> <div>Hg</div> <div>Mercury</div> <div>Switches</div> <div>Hg</div> <div>200.59</div> <div>Mercury is the only metal liquid at room temperature and</div> <div>Used in thermometers, dental fillings, and switches</div>

<div>81</div> <div>Screens</div> <div>Thallium</div> <div>204.38</div> <div>Thallium is 10 times more toxic than lead and was once sold</div> <div>Used in electronics and medical imaging</div>	<div>82</div> <div>Radiation</div> <div>Lead</div> <div>207.2</div> <div>Lead's toxicity may have contributed to the fall of Rome</div> <div>batteries, bullets, and radiation shielding</div>	<div>83</div> <div>Cosmetics</div> <div>Bismuth</div> <div>208.98</div> <div>Bismuth forms spectacular rainbow-colored oxide crystals</div> <div>Used in medicine and cosmetics</div>	<div>84</div> <div>Detectors</div> <div>Polonium</div> <div>209</div> <div>Polonium is 250 billion times more toxic than cyanide and</div> <div>Used in anticancer devices and neutron sources</div>
<div>85</div> <div>Therapy</div> <div>Astatine</div> <div>210</div> <div>Astatine is Earth's rarest element with less than 1 gram</div> <div>Used in medicine and scientific research</div>	<div>86</div> <div>Tracing</div> <div>Radon</div> <div>222</div> <div>Radon gas seepage causes 21,000 lung cancer deaths</div> <div>Used as tracer gas and in dating</div>	<div>87</div> <div>Experiments</div> <div>Francium</div> <div>223</div> <div>Francium is the most reactive metal with largest atomic</div> <div>Used in research and atomic clocks</div>	<div>88</div> <div>Treatment</div> <div>Radium</div> <div>226</div> <div>Radium was worth more than gold and glowed green due to</div> <div>Used in cancer treatment and luminous paints</div>
<div>89</div> <div>Lighting</div> <div>Actinium</div> <div>227</div> <div>Actinium glows blue-white in darkness and is 150 times</div> <div>Used in cancer treatment and neutron sources</div>	<div>90</div> <div>Mantles</div> <div>Thorium</div> <div>232.04</div> <div>Thorium is 3 times more abundant than uranium and</div> <div>Used in gas mantles and nuclear fuel</div>	<div>91</div> <div>Dating</div> <div>Protactinium</div> <div>231.04</div> <div>Protactinium costs \$280/gram making it one of the most</div> <div>Used in nuclear research and dating</div>	<div>92</div> <div>Weapons</div> <div>Uranium</div> <div>238.03</div> <div>Uranium-235's 1 gram releases energy equal to</div> <div>Used in nuclear fuel and weapons</div>
<div>93</div> <div>Detectors</div> <div>Neptunium</div> <div>237</div> <div>Neptunium was the first transuranium element created</div> <div>Used in smoke detectors and research</div>	<div>94</div> <div>Power</div> <div>Plutonium</div> <div>244</div> <div>Plutonium feels warm due to radioactive decay and is illegal</div> <div>Used in nuclear weapons and power</div>	<div>95</div> <div>Alarms</div> <div>Americium</div> <div>243</div> <div>Americium is the only man-made element available</div> <div>Used in smoke detectors and neutron sources</div>	<div>96</div> <div>Missions</div> <div>Curium</div> <div>247</div> <div>Curium glows purple-blue in darkness due to intense</div> <div>Used in research and space missions</div>
<div>97</div> <div>Cyclotron</div> <div>Berkelium</div> <div>247</div> <div>Berkelium was first synthesized at UC Berkeley</div> <div>Used in research and as electron source</div>	<div>98</div> <div>Starters</div> <div>Californium</div> <div>251</div> <div>Californium costs \$27 million per gram and is used to start</div> <div>Used in research and neutron sources</div>	<div>99</div> <div>Research</div> <div>Einsteinium</div> <div>252</div> <div>Einsteinium was discovered in hydrogen bomb debris from</div> <div>Used in research and medical applications</div>	<div>100</div> <div>Research</div> <div>Fermium</div> <div>257</div> <div>Fermium was found in H-bomb fallout like einsteinium and</div> <div>Used in research only</div>

<div> <div>101</div> <div>Research</div> <div> <div>258</div> <div>Used in research only</div> </div> </div> <div> <div>Md</div> <div>Mendelevium</div> <div>  </div> </div> <div> <div>Mendelevium was the first element created one atom at a time</div> </div>	<div> <div>102</div> <div>Research</div> <div> <div>259</div> <div>Used in research only</div> </div> </div> <div> <div>No</div> <div>Nobelium</div> <div>  </div> </div> <div> <div>Nobelium discovery was disputed for decades with the Soviet Union</div> </div>	<div> <div>103</div> <div>Research</div> <div> <div>262</div> <div>Used in research only</div> </div> </div> <div> <div>Lr</div> <div>Lawrencium</div> <div>  </div> </div> <div> <div>Lawrencium completes the actinide series and was the first element created one atom at a time</div> </div>	<div> <div>104</div> <div>Research</div> <div> <div>267</div> <div>Used in research only</div> </div> </div> <div> <div>Rf</div> <div>Rutherfordium</div> <div>  </div> </div> <div> <div>Rutherfordium was claimed by both Soviet and American scientists</div> </div>
<div> <div>105</div> <div>Research</div> <div> <div>270</div> <div>Used in research only</div> </div> </div> <div> <div>Db</div> <div>Dubnium</div> <div>  </div> </div> <div> <div>Dubnium was named after Dubna, Russia where Soviet scientists created it</div> </div>	<div> <div>106</div> <div>Research</div> <div> <div>271</div> <div>Used in research only</div> </div> </div> <div> <div>Sg</div> <div>Seaborgium</div> <div>  </div> </div> <div> <div>Seaborgium honors Glenn Seaborg, the only living person on the periodic table</div> </div>	<div> <div>107</div> <div>Research</div> <div> <div>270</div> <div>Used in research only</div> </div> </div> <div> <div>Bh</div> <div>Bohrium</div> <div>  </div> </div> <div> <div>Bohrium was named after Niels Bohr who developed quantum theory</div> </div>	<div> <div>108</div> <div>Research</div> <div> <div>277</div> <div>Used in research only</div> </div> </div> <div> <div>Hs</div> <div>Hassium</div> <div>  </div> </div> <div> <div>Hassium was named after Hesse, Germany where GSI is located</div> </div>
<div> <div>109</div> <div>Research</div> <div> <div>276</div> <div>Used in research only</div> </div> </div> <div> <div>Mt</div> <div>Meitnerium</div> <div>  </div> </div> <div> <div>Meitnerium honors Lise Meitner who discovered nuclear fission</div> </div>	<div> <div>110</div> <div>Research</div> <div> <div>285</div> <div>Used in research only</div> </div> </div> <div> <div>Ds</div> <div>Darmstadtium</div> <div>  </div> </div> <div> <div>Darmstadtium was named after Darmstadt, the latest city in Germany</div> </div>	<div> <div>111</div> <div>Research</div> <div> <div>282</div> <div>Used in research only</div> </div> </div> <div> <div>Rg</div> <div>Roentgenium</div> <div>  </div> </div> <div> <div>Roentgenium honors X-ray discoverer Wilhelm Röntgen</div> </div>	<div> <div>112</div> <div>Research</div> <div> <div>285</div> <div>Used in research only</div> </div> </div> <div> <div>Cn</div> <div>Copernicium</div> <div>  </div> </div> <div> <div>Copernicium was named after Copernicus who placed the heliocentric model of the universe</div> </div>
<div> <div>113</div> <div>Research</div> <div> <div>286</div> <div>Used in research only</div> </div> </div> <div> <div>Nh</div> <div>Nihonium</div> <div>  </div> </div> <div> <div>Nihonium was named after Japan (Nihon) where RIKEN is located</div> </div>	<div> <div>114</div> <div>Research</div> <div> <div>289</div> <div>Used in research only</div> </div> </div> <div> <div>Fl</div> <div>Flerovium</div> <div>  </div> </div> <div> <div>Flerovium honors Soviet physicist Flerov who founded the Joint Institute for Nuclear Research</div> </div>	<div> <div>115</div> <div>Research</div> <div> <div>290</div> <div>Used in research only</div> </div> </div> <div> <div>Mc</div> <div>Moscovium</div> <div>  </div> </div> <div> <div>Moscovium was named after Moscow where Russian scientists created it</div> </div>	<div> <div>116</div> <div>Research</div> <div> <div>293</div> <div>Used in research only</div> </div> </div> <div> <div>Lv</div> <div>Livermorium</div> <div>  </div> </div> <div> <div>Livermorium honors Lawrence Livermore Laboratory's 50th anniversary</div> </div>
<div> <div>117</div> <div>Research</div> <div> <div>294</div> <div>Used in research only</div> </div> </div> <div> <div>Ts</div> <div>Tennesine</div> <div>  </div> </div> <div> <div>Tennesine was named after Tennessee, discovered most recently</div> </div>	<div> <div>118</div> <div>Research</div> <div> <div>294</div> <div>Used in research only</div> </div> </div> <div> <div>Og</div> <div>Oganesson</div> <div>  </div> </div> <div> <div>Oganesson is the heaviest and most radioactive element with a half-life of 0.7 milliseconds</div> </div>		