



























































<div>Hydrogen</div> <div>1</div> <div>H</div> <div>Used as rocket fuel and in hydrogen cars</div> <div>Hydrogen is the lightest element and makes up 73.9% of the universe's visible matter</div> <div>1.008Fuel</div>	<div>Helium</div> <div>2</div> <div>He</div> <div>Used in balloons and as inert atmosphere for welding</div> <div>Helium has the lowest boiling point (-269°C) and becomes superfluid with zero viscosity</div> <div>4.0026Balloons</div>	<div>Lithium</div> <div>3</div> <div>Li</div> <div>Used in rechargeable batteries and mood stabilizers</div> <div>Lithium is the lightest metal that can float on water while violently reacting with it</div> <div>6.94Batteries</div>	<div>Beryllium</div> <div>4</div> <div>Be</div> <div>Used in aerospace alloys and nuclear reactors</div> <div>Beryllium is 6 times stronger than steel but weighs only 25% as much, yet highly toxic</div> <div>9.0122Alloys</div>
<div>Boron</div> <div>5</div> <div>B</div> <div>Used in ceramics, glass, and as neutron absorber</div> <div>Boron is the 5th element but essential for plants and harder than most metals when pure</div> <div>10.81Ceramics</div>	<div>Carbon</div> <div>6</div> <div>C</div> <div>Essential for all life and used in steel production</div> <div>Carbon is highest for 2 hundred million compounds, melting point of 3823K, and over 500 allotropes</div> <div>12.011Steel</div>	<div>Nitrogen</div> <div>7</div> <div>N</div> <div>Used in fertilizers and as liquid nitrogen coolant</div> <div>Nitrogen makes up 78% of Earth's atmosphere but is completely inert at room temperature</div> <div>14.007Fertilizers</div>	<div>Oxygen</div> <div>8</div> <div>O</div> <div>Essential for breathing and used in steel production</div> <div>Oxygen is Earth's most abundant element at 46% of crust mass and paramagnetic liquid</div> <div>15.999Breathing</div>
<div>Fluorine</div> <div>9</div> <div>F</div> <div>Used in toothpaste and water fluoridation</div> <div>Fluorine is the most reactive element that can corrode glass and concrete on contact</div> <div>18.998Toothpaste</div>	<div>Neon</div> <div>10</div> <div>Ne</div> <div>Used in neon signs and as inert gas in lighting</div> <div>Neon produces the most intense light discharge creating the classic orange-red glow</div> <div>20.18Signs</div>	<div>Sodium</div> <div>11</div> <div>Na</div> <div>Used in table salt and street lighting</div> <div>Sodium lamps are so efficient that one can outshine 100 incandescent bulbs combined</div> <div>22.99Salt</div>	<div>Magnesium</div> <div>12</div> <div>Mg</div> <div>Used in flares, alloys, and as dietary supplement</div> <div>Magnesium burns with 3000K white light so bright it can cause permanent eye damage</div> <div>24.305Flares</div>
<div>Aluminum</div> <div>13</div> <div>Al</div> <div>Used in beverage cans, foil, and aircraft parts</div> <div>Aluminum was worth more than gold until 1890s before efficient electrolytic extraction</div> <div>26.982Cans</div>	<div>Silicon</div> <div>14</div> <div>Si</div> <div>Used in computer chips, glass, and construction</div> <div>Silicon makes up 27% of Earth's crust and enabled the entire computer age revolution</div> <div>28.085Glass</div>	<div>Phosphorus</div> <div>15</div> <div>P</div> <div>Used in fertilizers, matches, and DNA structure</div> <div>White phosphorus glows green in darkness but is essential for life despite being toxic</div> <div>30.974Matches</div>	<div>Sulfur</div> <div>16</div> <div>S</div> <div>Used in rubber vulcanization and gunpowder</div> <div>Sulfur is second for forming 30+ allotropes and creates yellow crystals in volcanic regions</div> <div>32.06Rubber</div>
<div>Chlorine</div> <div>17</div> <div>Cl</div> <div>Used in pool disinfection and PVC production</div> <div>Chlorine has the highest electron affinity (349 kJ/mol) and was WWI's first poison gas</div> <div>35.45Pools</div>	<div>Argon</div> <div>18</div> <div>Ar</div> <div>Used in welding and incandescent light bulbs</div> <div>Argon was Earth's first isolated noble gas and makes up nearly 1% of our atmosphere</div> <div>39.948Welding</div>	<div>Potassium</div> <div>19</div> <div>K</div> <div>Used in fertilizers and soap production</div> <div>Potassium is so violently reactive it ignites spontaneously and must be stored in oil</div> <div>39.098Fertilizer</div>	<div>Calcium</div> <div>20</div> <div>Ca</div> <div>Used in bones, teeth, and concrete production</div> <div>Calcium phosphate comprises 70% of bone mass giving vertebrates their rigid structure</div> <div>40.078Bones</div>

<div>Scandium</div> <div> <div>21</div> <div>  <div>Sc</div> </div> <div>Used in aerospace alloys and baseball bats</div> <div>Scandium is paradoxically rarer than many 'rare earth' elements despite being lighter</div> <div>44.956Aerospace</div> </div>	<div>Titanium</div> <div> <div>22</div> <div>  <div>Ti</div> </div> <div>Used in aircraft, medical implants, and paints</div> <div>Titanium has the highest strength-to-weight ratio of all metals at 45% lighter than steel</div> <div>47.867Implants</div> </div>	<div>Vanadium</div> <div> <div>23</div> <div>  <div>V</div> </div> <div>Used in steel alloys and as catalyst</div> <div>Vanadium can exist in 5 different oxidation states creating rainbow-colored solutions</div> <div>50.942Steel</div> </div>	<div>Chromium</div> <div> <div>24</div> <div>  <div>Cr</div> </div> <div>Used in stainless steel and chrome plating</div> <div>Chromium gives rubies red and emeralds green color while being the hardest pure metal</div> <div>51.996Stainless</div> </div>
<div>Manganese</div> <div> <div>25</div> <div>  <div>Mn</div> </div> <div>Used in steel production and battery electrodes</div> <div>Manganese nodules carpet vast ocean floors containing trillions of tons of the element</div> <div>54.938Batteries</div> </div>	<div>Iron</div> <div> <div>26</div> <div>  <div>Fe</div> </div> <div>Used in construction, tools, and magnets</div> <div>Iron comprises 32.1% of Earth's total mass with most concentrated in the molten core</div> <div>55.845Magnets</div> </div>	<div>Cobalt</div> <div> <div>27</div> <div>  <div>Co</div> </div> <div>Used in magnets, catalysts, and blue glass</div> <div>Cobalt blue glass has been prized for 4000 years and retains color at 1000°C heat</div> <div>58.933Motors</div> </div>	<div>Nickel</div> <div> <div>28</div> <div>  <div>Ni</div> </div> <div>Used in coins, batteries, and stainless steel</div> <div>Nickel-iron meteorites delivered much of Earth's accessible nickel from space impacts</div> <div>58.693Coins</div> </div>
<div>Copper</div> <div> <div>29</div> <div>  <div>Cu</div> </div> <div>Used in electrical wiring and plumbing pipes</div> <div>Copper naturally kills bacteria and viruses within hours making it self-sterilizing</div> <div>63.546Wire</div> </div>	<div>Zinc</div> <div> <div>30</div> <div>  <div>Zn</div> </div> <div>Used in galvanizing steel and brass alloys</div> <div>Zinc deficiency causes loss of taste/smell and affects 2 billion people worldwide</div> <div>65.38Galvanizing</div> </div>	<div>Gallium</div> <div> <div>31</div> <div>  <div>Ga</div> </div> <div>Used in semiconductors and LEDs</div> <div>Gallium melts at 29.8°C in hand temperature but boils at 2400°C with the widest liquid range</div> <div>69.723Electronics</div> </div>	<div>Germanium</div> <div> <div>32</div> <div>  <div>Ge</div> </div> <div>Used in fiber optics and transistors</div> <div>Germanium was predicted by Mendeleev 15 years before discovery with exact properties</div> <div>72.63Semiconductors</div> </div>
<div>Arsenic</div> <div> <div>33</div> <div>  <div>As</div> </div> <div>Used in wood preservatives and semiconductors</div> <div>Arsenic has been humanity's poison of choice for over 2000 years earning 'King of Poisons'</div> <div>74.922Pesticides</div> </div>	<div>Selenium</div> <div> <div>34</div> <div>  <div>Se</div> </div> <div>Used in photocopying and glass coloring</div> <div>Selenium deficiency causes fatal white muscle disease and is toxic in excess amounts</div> <div>78.971Glass</div> </div>	<div>Bromine</div> <div> <div>35</div> <div>  <div>Br</div> </div> <div>Used as antiseptic and in flame retardants</div> <div>Bromine is the only liquid non-metal but it evaporates quickly from 1 mL to 3 liters of toxic gas</div> <div>79.904Antiseptic</div> </div>	<div>Krypton</div> <div> <div>36</div> <div>  <div>Kr</div> </div> <div>Used in energy-efficient windows and lasers</div> <div>Krypton was used in ultra-bright airport runway lighting systems and old camera flashes</div> <div>83.798Lasers</div> </div>
<div>Rubidium</div> <div> <div>37</div> <div>  <div>Rb</div> </div> <div>Used in atomic clocks and medical tracers</div> <div>Rubidium ignites spontaneously in air and was used in early vacuum tubes for electronics</div> <div>85.468Atomic</div> </div>	<div>Strontium</div> <div> <div>38</div> <div>  <div>Sr</div> </div> <div>Used in fireworks and flares for red color</div> <div>Strontium-90 fallout creates the brilliant red in fireworks but is dangerously radioactive</div> <div>87.62Fireworks</div> </div>	<div>Yttrium</div> <div> <div>39</div> <div>  <div>Y</div> </div> <div>Used in lasers and as cancer treatment</div> <div>Yttrium with barium carbon oxide named YBCO makes the highest temperature superconductors at 92K</div> <div>88.906Catalysts</div> </div>	<div>Zirconium</div> <div> <div>40</div> <div>  <div>Zr</div> </div> <div>Used in nuclear reactors and ceramics</div> <div>Zirconium is virtually immune to corrosion up to 1270K and used in nuclear reactors</div> <div>91.224Jet</div> </div>

<div> <div>Niobium</div> <div> <div>41</div> <div>Nb</div> <div>Used in jet engines and MRI scanners</div> <div> <i>Niobium is superconducting below 9K and was originally called columbium in America</i> </div> <div> <div>92.906</div> <div>Steel</div> </div> </div> </div>	<div> <div>Molybdenum</div> <div> <div>42</div> <div>Mo</div> <div>Used in steel alloys and high-temp lubricants</div> <div> <i>Molybdenum has the 6th highest melting point at 2896K and strengthens steel dramatically</i> </div> <div> <div>95.95</div> <div>Lubricants</div> </div> </div> </div>	<div> <div>Technetium</div> <div> <div>43</div> <div>Tc</div> <div>Used in medical imaging and as tracer</div> <div> <i>Technetium was the first artificially created element filling Mendeleev's predicted gap</i> </div> <div> <div>98</div> <div>Medicine</div> </div> </div> </div>	<div> <div>Ruthenium</div> <div> <div>44</div> <div>Ru</div> <div>Used in electrical contacts and hard disks</div> <div> <i>Ruthenium is the scarcest platinum group metal and costs \$1600 per troy ounce</i> </div> <div> <div>101.07</div> <div>Electronics</div> </div> </div> </div>
<div> <div>Rhodium</div> <div> <div>45</div> <div>Rh</div> <div>Used in catalytic converters and jewelry</div> <div> <i>Rhodium is the most expensive precious metal at \$14,000+ per ounce, rarer than gold</i> </div> <div> <div>102.91</div> <div>Catalysts</div> </div> </div> </div>	<div> <div>Palladium</div> <div> <div>46</div> <div>Pd</div> <div>Used in catalytic converters and dentistry</div> <div> <i>Palladium can absorb 900 times its volume in hydrogen like a metallic sponge</i> </div> <div> <div>106.42</div> <div>Jewelry</div> </div> </div> </div>	<div> <div>Silver</div> <div> <div>47</div> <div>Ag</div> <div>Used in jewelry, mirrors, and photography</div> <div> <i>Silver has the highest electrical conductivity of all elements at room temperature</i> </div> <div> <div>107.87</div> <div>Mirrors</div> </div> </div> </div>	<div> <div>Cadmium</div> <div> <div>48</div> <div>Cd</div> <div>Used in batteries, pigments, and solar panels</div> <div> <i>Cadmium red paint was banned after causing severe poisoning in artists for decades</i> </div> <div> <div>112.41</div> <div>Batteries</div> </div> </div> </div>
<div> <div>Indium</div> <div> <div>49</div> <div>In</div> <div>Used in semiconductors and LCD screens</div> <div> <i>Indium is softer than lead and can be scratched with a fingernail despite being metal</i> </div> <div> <div>114.82</div> <div>Semiconductors</div> </div> </div> </div>	<div> <div>Tin</div> <div> <div>50</div> <div>Sn</div> <div>Used in solder, cans, and bronze alloys</div> <div> <i>Tin produces a distinctive 'tin cry' scream when bent due to crystal twinning</i> </div> <div> <div>118.71</div> <div>Solder</div> </div> </div> </div>	<div> <div>Antimony</div> <div> <div>51</div> <div>Sb</div> <div>Used in flame retardants and semiconductors</div> <div> <i>Fluoroantimonic acid is 10 quintillion times stronger than sulfuric acid - the strongest known</i> </div> <div> <div>121.76</div> <div>Flame</div> </div> </div> </div>	<div> <div>Tellurium</div> <div> <div>52</div> <div>Te</div> <div>Used in solar panels and rubber vulcanization</div> <div> <i>Tellurium-128 has the longest known half-life at 2.2 septillion years - nearly stable</i> </div> <div> <div>127.6</div> <div>Solar</div> </div> </div> </div>
<div> <div>Iodine</div> <div> <div>53</div> <div>I</div> <div>Used as antiseptic and in photography</div> <div> <i>Iodine deficiency affects 2 billion people causing goiter and developmental disability</i> </div> <div> <div>126.9</div> <div>Antiseptic</div> </div> </div> </div>	<div> <div>Xenon</div> <div> <div>54</div> <div>Xe</div> <div>Used in ion drives and medical anesthesia</div> <div> <i>Xenon is the rarest gas with 90 grams per million kilograms of air</i> </div> <div> <div>131.29</div> <div>Anesthesia</div> </div> </div> </div>	<div> <div>Cesium</div> <div> <div>55</div> <div>Cs</div> <div>Used in atomic clocks and oil drilling</div> <div> <i>Caesium is the softest metal and its hydroxide is the strongest base ever discovered</i> </div> <div> <div>132.91</div> <div>Atomic</div> </div> </div> </div>	<div> <div>Barium</div> <div> <div>56</div> <div>Ba</div> <div>Used in X-ray imaging and drilling fluids</div> <div> <i>Barium compounds create brilliant green fireworks but are lethally toxic if ingested</i> </div> <div> <div>137.33</div> <div>X-rays</div> </div> </div> </div>
<div> <div>Lanthanum</div> <div> <div>57</div> <div>La</div> <div>Used in lighter flints and camera lenses</div> <div> <i>Lanthanum remained undiscovered in 'pure' cerium samples for 83 years of confusion</i> </div> <div> <div>138.91</div> <div>Lighter</div> </div> </div> </div>	<div> <div>Cerium</div> <div> <div>58</div> <div>Ce</div> <div>Used in catalysts and glass polishing</div> <div> <i>Cerium is the most abundant rare earth comprising 0.006% of Earth's crust mass</i> </div> <div> <div>140.12</div> <div>Catalysts</div> </div> </div> </div>	<div> <div>Praseodymium</div> <div> <div>59</div> <div>Pr</div> <div>Used in aircraft engines and magnets</div> <div> <i>Praseodymium means 'green twin' creating emerald-green compounds and yellow metal</i> </div> <div> <div>140.91</div> <div>Magnets</div> </div> </div> </div>	<div> <div>Neodymium</div> <div> <div>60</div> <div>Nd</div> <div>Used in powerful permanent magnets</div> <div> <i>Neodymium creates the strongest permanent magnets lifting 1000 times their own weight</i> </div> <div> <div>144.24</div> <div>Magnets</div> </div> </div> </div>

<div>Promethium</div> <div> <div>61</div> <div>  <div>Pm</div> </div> <div>Used in nuclear batteries and research</div> <div>Promethium is the only radioactive rare earth and powers space missions for decades</div> <div>145Batteries</div> </div>	<div>Samarium</div> <div> <div>62</div> <div>  <div>Sm</div> </div> <div>Used in magnets and cancer treatment</div> <div>Samarium magnets work at 350°C and have the highest neutron absorption cross-section</div> <div>150.36Magnets</div> </div>	<div>Europium</div> <div> <div>63</div> <div>  <div>Eu</div> </div> <div>Used in red phosphors for TV screens</div> <div>Europium is the softest rare earth and the most reactive, tarnishing rapidly in air</div> <div>151.96Phosphors</div> </div>	<div>Gadolinium</div> <div> <div>64</div> <div>  <div>Gd</div> </div> <div>Used in MRI contrast agents and neutron capture</div> <div>Gadolinium has the highest magnetic moment and is used in MRI contrast enhancement</div> <div>157.25MRI</div> </div>
<div>Terbium</div> <div> <div>65</div> <div>  <div>Tb</div> </div> <div>Used in green phosphors and magnets</div> <div>Terbium glows intense green under UV and is essential for energy-efficient lighting</div> <div>158.93Magnets</div> </div>	<div>Dysprosium</div> <div> <div>66</div> <div>  <div>Dy</div> </div> <div>Used in lasers and hard disk drives</div> <div>Dysprosium becomes strongly magnetic only below -180°C with highest magnetic strength</div> <div>162.5Lasers</div> </div>	<div>Holmium</div> <div> <div>67</div> <div>  <div>Ho</div> </div> <div>Used in magnets and medical devices</div> <div>Holmium possesses the strongest magnetic field of any element at 4.5 Tesla saturation</div> <div>164.93Magnets</div> </div>	<div>Erbium</div> <div> <div>68</div> <div>  <div>Er</div> </div> <div>Used in fiber optic amplifiers and lasers</div> <div>Erbium amplifies light in fiber optic cables enabling global internet communications</div> <div>167.26Fiber</div> </div>
<div>Thulium</div> <div> <div>69</div> <div>  <div>Tm</div> </div> <div>Used in X-ray sources and portable equipment</div> <div>Thulium is the least abundant rare earth metal and possibly the most useless natural element</div> <div>168.93X-rays</div> </div>	<div>Ytterbium</div> <div> <div>70</div> <div>  <div>Yb</div> </div> <div>Used in lasers and stress gauges</div> <div>Ytterbium expands 26% during phase transition and is used in atomic clocks</div> <div>173.05Lasers</div> </div>	<div>Lutetium</div> <div> <div>71</div> <div>  <div>Lu</div> </div> <div>Used in catalysts and medical imaging</div> <div>Lutetium is the hardest, densest rare earth and was the last lanthanide discovered</div> <div>174.97Catalysts</div> </div>	<div>Hafnium</div> <div> <div>72</div> <div>  <div>Hf</div> </div> <div>Used in tungsten carbide and nuclear reactors</div> <div>Hafnium has nearly identical properties to zirconium due to lanthanide contraction</div> <div>178.49Carbide</div> </div>
<div>Tantalum</div> <div> <div>73</div> <div>  <div>Ta</div> </div> <div>Used in electronics and surgical instruments</div> <div>Tantalum is virtually immune to all acids except hydrofluoric at high temperatures</div> <div>180.95Electronics</div> </div>	<div>Tungsten</div> <div> <div>74</div> <div>  <div>W</div> </div> <div>Used in light bulb filaments and X-ray tubes</div> <div>Tungsten has the highest melting point at 3695K and tensile strength of all metals</div> <div>183.84Bulbs</div> </div>	<div>Rhenium</div> <div> <div>75</div> <div>  <div>Re</div> </div> <div>Used in catalysts and jet engine parts</div> <div>Rhenium has the highest boiling point at 5869K and is the last stable element found</div> <div>186.21Catalysts</div> </div>	<div>Osmium</div> <div> <div>76</div> <div>  <div>Os</div> </div> <div>Used in fountain pen tips and electrical contacts</div> <div>Osmium is the densest element at 22.6 g/cm³ and costs \$400 per troy ounce</div> <div>190.23Fountain</div> </div>
<div>Iridium</div> <div> <div>77</div> <div>  <div>Ir</div> </div> <div>Used in spark plugs and cancer treatment</div> <div>Iridium is the most corrosion-resistant element and 2nd densest element at 22.42 g/cm³</div> <div>192.22Catalysts</div> </div>	<div>Platinum</div> <div> <div>78</div> <div>  <div>Pt</div> </div> <div>Used in jewelry, catalysts, and electronics</div> <div>Platinum is 30 times rarer than gold and catalyzes 20% of all chemical processes</div> <div>195.08Jewelry</div> </div>	<div>Gold</div> <div> <div>79</div> <div>  <div>Au</div> </div> <div>Used in jewelry, electronics, and dentistry</div> <div>Gold is so chemically inert it never tarnishes and has been treasured for 6000 years</div> <div>196.97Electronics</div> </div>	<div>Mercury</div> <div> <div>80</div> <div>  <div>Hg</div> </div> <div>Used in thermometers, dental fillings, and switches</div> <div>Mercury is the only metal liquid at room temperature and expands linearly with heat</div> <div>200.59Thermometers</div> </div>

<div>Thallium</div> <div>81</div> <div></div> <div>Tl</div> <div>Used in electronics and medical imaging</div> <div>Thallium is 10 times more toxic than lead and was once sold as rat poison</div> <div>204.38</div> <div>Electronics</div>	<div>Lead</div> <div>82</div> <div></div> <div>Pb</div> <div>Used in car batteries, bullets, and radiation shielding</div> <div>Lead's toxicity may have contributed to the fall of Rome through poisoned water pipes</div> <div>207.2</div> <div>Batteries</div>	<div>Bismuth</div> <div>83</div> <div></div> <div>Bi</div> <div>Used in medicine and cosmetics</div> <div>Bismuth forms spectacular rainbow-colored oxide crystals and expands when solidifying</div> <div>208.98</div> <div>Medicine</div>	<div>Polonium</div> <div>84</div> <div></div> <div>Po</div> <div>Used in antistatic devices and neutron sources</div> <div>Polonium is 250 billion times more toxic than cyanide and the most radioactive natural element</div> <div>209</div> <div>Detectors</div>
<div>Astatine</div> <div>85</div> <div></div> <div>At</div> <div>Used in medicine and scientific research</div> <div>Astatine is Earth's rarest element with less than 1 gram existing at any time</div> <div>210</div> <div>Medicine</div>	<div>Radon</div> <div>86</div> <div></div> <div>Rn</div> <div>Used as tracer gas and in dating</div> <div>Radon gas seepage causes 21,000 lung cancer deaths annually in the US alone</div> <div>222</div> <div>Gas</div>	<div>Francium</div> <div>87</div> <div></div> <div>Fr</div> <div>Used in research and atomic clocks</div> <div>Francium is the most reactive metal with largest atomic radius and shortest half-life of 22 minutes</div> <div>223</div> <div>Research</div>	<div>Radium</div> <div>88</div> <div></div> <div>Ra</div> <div>Used in cancer treatment and luminous paints</div> <div>Radium was worth more than gold and glowed green due to intense radioactive decay</div> <div>226</div> <div>Medicine</div>
<div>Actinium</div> <div>89</div> <div></div> <div>Ac</div> <div>Used in cancer treatment and neutron sources</div> <div>Actinium glows blue-white in darkness and is 150 times more radioactive than radium</div> <div>227</div> <div>Medicine</div>	<div>Thorium</div> <div>90</div> <div></div> <div>Th</div> <div>Used in gas mantles and nuclear fuel</div> <div>Thorium is 3 times more abundant than uranium and could power civilization for millennia</div> <div>232.04</div> <div>Gas</div>	<div>Protactinium</div> <div>91</div> <div></div> <div>Pa</div> <div>Used in nuclear research and dating</div> <div>Protactinium costs \$280/gram making it one of the most expensive elements to obtain</div> <div>231.04</div> <div>Nuclear</div>	<div>Uranium</div> <div>92</div> <div></div> <div>U</div> <div>Used in nuclear fuel and weapons</div> <div>Uranium-235's 1 gram releases energy equal to burning 3 tons of coal completely</div> <div>238.03</div> <div>Fuel</div>
<div>Neptunium</div> <div>93</div> <div></div> <div>Np</div> <div>Used in smoke detectors and research</div> <div>Neptunium was the first transuranium element created and is named after planet Neptune</div> <div>237</div> <div>Detectors</div>	<div>Plutonium</div> <div>94</div> <div></div> <div>Pu</div> <div>Used in nuclear weapons and power</div> <div>Plutonium feels warm due to radioactive decay and is illegal for civilians to possess</div> <div>244</div> <div>Weapons</div>	<div>Americium</div> <div>95</div> <div></div> <div>Am</div> <div>Used in smoke detectors and neutron sources</div> <div>Americium is the only man-made element available to the public in stores</div> <div>243</div> <div>Detectors</div>	<div>Curium</div> <div>96</div> <div></div> <div>Cm</div> <div>Used in research and space missions</div> <div>Curium glows purple-blue in darkness due to intense radioactivity and rapid decay</div> <div>247</div> <div>Research</div>
<div>Berkelium</div> <div>97</div> <div></div> <div>Bk</div> <div>Used in research and as electron source</div> <div>Berkelium was first synthesized at UC Berkeley using the 60-inch cyclotron in 1949</div> <div>247</div> <div>Research</div>	<div>Californium</div> <div>98</div> <div></div> <div>Cf</div> <div>Used in research and neutron sources</div> <div>Californium costs \$27 million per gram and is used to start nuclear reactors</div> <div>251</div> <div>Research</div>	<div>Einsteinium</div> <div>99</div> <div></div> <div>Es</div> <div>Used in research and medical applications</div> <div>Einsteinium was discovered in hydrogen bomb debris from the first H-bomb test in 1952</div> <div>252</div> <div>Research</div>	<div>Fermium</div> <div>100</div> <div></div> <div>Fm</div> <div>Used in research only</div> <div>Fermium was found in H-bomb fallout like einsteinium and named after physicist Enrico Fermi</div> <div>257</div> <div>Research</div>

Mendelevium	Nobelium	Lawrencium	Rutherfordium
<div>101</div> <div></div> <div>Md</div> <div>Used in research only</div> <div>Mendelevium was the first element created one atom at a time using particle accelerators</div> <div>258Research</div>	<div>102</div> <div></div> <div>No</div> <div>Used in research only</div> <div>Nobelium discovery was disputed for decades with Soviet, American, and Swedish claims</div> <div>259Research</div>	<div>103</div> <div></div> <div>Lr</div> <div>Used in research only</div> <div>Lawrencium completes the actinide series and was synthesized at Berkeley in 1961</div> <div>262Research</div>	<div>104</div> <div></div> <div>Rf</div> <div>Used in research only</div> <div>Rutherfordium was claimed by both Soviet and American teams causing naming disputes</div> <div>267Research</div>
Dubnium	Seaborgium	Bohrium	Hassium
<div>105</div> <div></div> <div>Db</div> <div>Used in research only</div> <div>Dubnium was named after Dubna, Russia where Soviet scientists first claimed discovery</div> <div>270Research</div>	<div>106</div> <div></div> <div>Sg</div> <div>Used in research only</div> <div>Seaborgium honors Glenn Seaborg, the only living person to have an element named for them</div> <div>271Research</div>	<div>107</div> <div></div> <div>Bh</div> <div>Used in research only</div> <div>Bohrium was named after Niels Bohr who developed quantum mechanical model of atoms</div> <div>270Research</div>	<div>108</div> <div></div> <div>Hs</div> <div>Used in research only</div> <div>Hassium was named after Hesse, Germany where GSI laboratory first synthesized it</div> <div>277Research</div>
Meitnerium	Darmstadtium	Roentgenium	Copernicium
<div>109</div> <div></div> <div>Mt</div> <div>Used in research only</div> <div>Meitnerium honors Lise Meitner who discovered nuclear fission</div> <div>276Research</div>	<div>110</div> <div></div> <div>Ds</div> <div>Used in research only</div> <div>Darmstadtium was named after Darmstadt, the latest city to receive elemental recognition</div> <div>281Research</div>	<div>111</div> <div></div> <div>Rg</div> <div>Used in research only</div> <div>Roentgenium honors X-ray discoverer Wilhelm Röntgen though it doesn't emit X-rays</div> <div>282Research</div>	<div>112</div> <div></div> <div>Cn</div> <div>Used in research only</div> <div>Copernicium was named after Copernicus who placed the Sun at the solar system's center</div> <div>285Research</div>
Nihonium	Flerovium	Moscovium	Livermorium
<div>113</div> <div></div> <div>Nh</div> <div>Used in research only</div> <div>Nihonium was named after Japan (Nihon) where RIKEN laboratory first synthesized it in 2004</div> <div>286Research</div>	<div>114</div> <div></div> <div>Fl</div> <div>Used in research only</div> <div>Flerovium honors Soviet physicist Flerov who founded heavy element research in USSR</div> <div>289Research</div>	<div>115</div> <div></div> <div>Mc</div> <div>Used in research only</div> <div>Moscovium was named after Moscow where Russian scientists contributed to superheavy research</div> <div>290Research</div>	<div>116</div> <div></div> <div>Lv</div> <div>Used in research only</div> <div>Livermorium honors Lawrence Livermore Laboratory's contributions to superheavy elements</div> <div>293Research</div>
Tennessine	Oganesson		
<div>117</div> <div></div> <div>Ts</div> <div>Used in research only</div> <div>Tennessine was named after Tennessee, discovered most recently in 2010 at Oak Ridge</div> <div>294Research</div>	<div>118</div> <div></div> <div>Og</div> <div>Used in research only</div> <div>Oganesson is the heaviest and most radioactive element with the shortest 0.7ms half-life</div> <div>294Research</div>		