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| Both the forward and backwards reactions are happening at the same time |
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| They are equal |
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| When a system at equilibrium is disturbed, the position of equilibrium will shift to reduce the disturbance |
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| Concentrations of the reactants and products remain constant  The rate of the forward reaction is the same as the rate of the backwards reaction |
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| The catalyst speeds up both the forward and backwards reaction  The increase in rate is the same |
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| The equilibrium constant depends on the temperature |
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| Explain why the equilibrium constant might have no units |
| If the number of moles on both sides was equal |
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| For a system where the forward reaction is exothermic, why might a compromise temperature be used? |
| High temperature will give low yield  Low temperature will give a slow reaction |
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| Forward reaction is exothermic |
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