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Chapter 23 Nuclear Chemistry Notes 1 CHAPTER 23 NUCLEAR CHEMISTRY 23.1 THE NATURE OF NUCLEAR REACTIONS radioactivity - the spontaneous decay of an unstable nucleus with accompanying emission of radiation. nuclide - atom with a specific number of protons and neutrons in its nucleus. ⇒ There are 271 stable nuclides in nature, others are radioactive

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chapter 23: nuclear chemistry 23.27 A truly first-order rate law implies that the mechanism is unimolecular; in other words the rate is determined only by the properties of the decaying atom or molecule and does not depend on collisions or interactions with other objects.

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Chapter 23: Nuclear Chemistry 23.1: The Nature of Nuclear Reactions Nucleons: - the particles that make up a nucleus of an atom (protons, (1 1 p + or 1 1 H) and neutrons, (1 0 n)). Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

Chapter 23 Nuclear Chemistry Notes (answers)

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Chapter 22 Review: Nuclear Chemistry. History Just because it is interesting. Marie Curie: Henri Becquerel (Physics professor, 1852 - 1908) discovered radioactivity in 1896 by placing a photographic plate near a box of uranium ore. When the plate "developed" on its own, he knew the ore contained something giving off energy. ... Never used in ...

Chapter 22 Review: Nuclear Chemistry

In this lecture I'll teach you about nuclear chemistry. I'll first show you how to determine an element's number of protons, electrons, and neutrons from its atomic symbol.

Chapter 21 - Nuclear Chemistry: Part 1 of 9

Chapter 21: Nuclear Chemistry Exam . 20 Questions | By Jcodylewis | Last updated: Jan 29, 2013 . Please take the quiz to rate it. Title of New (Duplicated) Quiz: Duplicate Quiz Cancel+ Success! A copy of this quiz is in your dashboard. Go to My Dashboard This quiz is in draft mode and is not available yet.

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Chapter 22 Review Nuclear Chemistry Section 1 Answers

CHAPTER 23 NUCLEAR CHEMISTRY 23.5 (a) The atomic number sum and the mass number sum must remain the same on both sides of a nuclear equation. On the left side of this equation the atomic number sum is 13 (12 + 1) and the mass number sum is 27 (26 + 1).

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