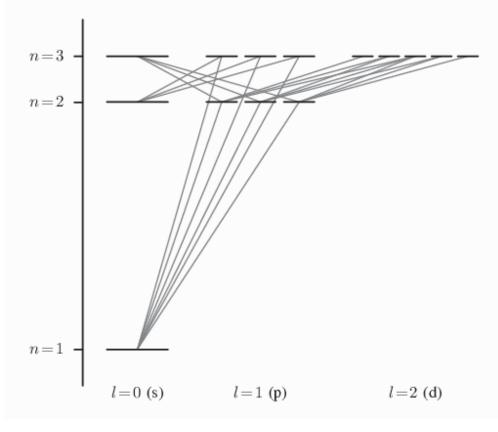
PROBLEM STATEMENT

Selection rules

The hydrogen atom consists of a single electron orbiting a positively charged nucleus. The electron can exist only in discrete orbitals characterized by the radial quantum number $n=1,2,3,\ldots$ and the angular momentum quantum numbers $l=0,1,2,\ldots,n-1$ and $m=-l,\ldots,-1,0,1,\ldots,l$. The energy of each orbital, $E_n=-(13.6\,\text{eV})/n^2$, is a function of the radial quantum number alone. Hence, each energy level is g_n -fold degenerate, where $g_n=\sum_{l=0}^{n-1}(2l+1)=n^2$. (In other words, there are $g_n=n^2$ states having the same energy E_n .)



Transitions between orbitals can occur if the electron absorbs or emits a photon. But since a quantum of light has intrinsic angular momentum (1 in units of \hbar), conservation laws put a strict limit on which atomic transitions are possible. This leads to the famous electric dipole *selection rules*: $\Delta I = \pm 1$ and $\Delta m = 0, \pm 1$. Allowed transitions between the various states with n = 1, 2, 3 are shown in the following diagram.

Complete the program selection.cpp so that it computes all possible transitions of the form $n_2 \rightarrow n_1$ with $n_1 < n_2 \le 20$. Have the program output the results in four columns indicating the initial and final radial quantum numbers, the number of allowed pathways, the energy of the emitted photon $\Delta E = E_2 - E_1$, and its wavelength $\lambda = hc/\Delta E$ in nanometers. (Recall that $hc = 1240 \, eV \cdot nm$.)

```
$ make selection
g++ -o selection selection.cpp -O2 -ansi -pedantic -Wall -lm
$ ./selection | head -n5
2->1 3
                             121.569
                   10.2
          3
3->1
                 12.0889
                              102.574
3->2
        15
                 1.88889
                              656.471
4->1 3
4->2 15
         3
                   12.75
                              97.2549
                    2.55
                               486.275
```

Modify selection.cpp so that the spectral lines in the visible spectrum, 380 nm $< \lambda <$ 750 nm, are marked with a lowercase v .

```
$ make selection
g++ -o selection selection.cpp -O2 -ansi -pedantic -Wall -lm
$ ./selection | head -n5
                   10.2
2->1 3
                             121.569
          3
3->1
                 12.0889
                             102.574
3->2v
         15
                 1.88889
                             656,471
4->1
          3
                   12.75
                              97.2549
4->2v 15
                   2.55
                              486.275
$ ./selection | grep v
3->2v 15
                1.88889
                             656.471
4->2v
          15
                   2.55
                              486.275
5->2v
         15
                   2.856
                             434.174
         15
                 3.02222
6->2v
                             410.294
7->2v
          15
                 3.12245
                               397.124
8->2v
          15
                  3.1875
                               389.02
9->2v
          15
                  3.2321
                               383.652
```