

Calculation:

Weight of copper sulphate pentahydrate present in 100 cm<sup>3</sup> of the given solution = 55 mg

Weight of CuSO<sub>4</sub> · 5H<sub>2</sub>O in 1 cm<sup>3</sup> of its solution =  $\frac{55}{100}$  mg

Weight of copper present in 1 cm<sup>3</sup> of its solution =  $\frac{55 \times 63.54}{249.54}$  mg

Model Procedure Flow Chart:

Filim, 46 loc Cop slphe S. lsbnsi,  
S<sup>3</sup> volumeda Flst Bede. Ad a  
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by i diklledwa. PufR a blanc salbirn  
2Se omorin di-les inze Pls FI  
Mix th\* 5 ns pp  
6- nm Bltes Colimeka  
Likaly t blane slbion mate i3e  
Me di e e h andd f dHh  
m eloda 2ph  
Fam 2K Brd  
Model graph:

nLan vol C Sl

Model Calculation:

$$\frac{\text{Wt of Cu in 'ren-ss}}{\text{Wt of Cu in 'ren-ss}} = \frac{\text{Vol of CuSO}_4 \text{ in cm}^3}{\text{Vol of CuSO}_4 \text{ in cm}^3} \times \frac{\text{Wt. of Copper (mg)}}{\text{Wt. of Copper (mg)}}$$

$$\frac{\text{Wt of Cu in 'ren-ss}}{\text{Wt of Cu in 'ren-ss}} = \frac{\text{Vol of CuSO}_4 \text{ in cm}^3}{\text{Vol of CuSO}_4 \text{ in cm}^3} \times \frac{\text{Wt. of Copper (mg)}}{\text{Wt. of Copper (mg)}}$$

Tabulation:

Sl. No.	Vol of CuSO <sub>4</sub> in cm	Optical Density	Wt. of Copper (mg)
1	0 [ Blank]		
2.	2		0.2
3.	4	0.10	0.2
4.	6	0.18	0.2
5.	8		(.12
	10		1.4

Unknown-1 0.2

Calculation:

$$\frac{\text{Wt of Cu in 'ren-ss}}{\text{Wt of Cu in 'ren-ss}} = \frac{\text{Vol of CuSO}_4 \text{ in cm}^3}{\text{Vol of CuSO}_4 \text{ in cm}^3} \times \frac{\text{Wt. of Copper (mg)}}{\text{Wt. of Copper (mg)}}$$

9Va

WE i:cu:uxo:/4eOS6~

$$\frac{\text{Cu} \times 0.1}{\text{Cu} \times 0.1}$$

pwt C lox1

42c

wt of Cu in sludki: 9yo1 0:692m

OAS

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Inference:

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Relevance to Society & Environment:

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Report:

1. Volume of unknown solution . . cm
2. Amount of Copper in the given unknown solution . mg

Evaluation of experiment -8

Components		Marks	
		Max	Obtained
Model Procedure,		16	2019
Model Graph &			
Calculation			
Expected Volume &			
Execution		0L 03	
Inference & Societal			
Relevance			

Total

Signature of Teacher